Atmospheric Oxidation Mechanism and Kinetics of Indole Initiated by ·OH and ·Cl: A Computational Study

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Abstract. The atmospheric chemistry of organic nitrogen compounds (ONCs) is of great importance for understanding the formation of carcinogenic nitrosamines, and ONC oxidation products might influence atmospheric aerosol particle formation and growth. Indole is a polyfunctional heterocyclic secondary amine with global emission quantity almost equivalent to that of trimethylamine, the amine with the highest atmospheric emission. However, the atmospheric chemistry of indole remains unclear. Herein, the reactions of indole with ·OH/·Cl, and subsequent reactions of resulting indole-radicals with O_2 under 200 ppt NO and 50 ppt HO_2· conditions, were investigated by a combination of quantum chemical calculations and kinetics modeling. The results indicate that ·OH addition is the dominant pathway for the reaction of ·OH with indole. However, both ·Cl addition and H-abstraction are feasible for the corresponding reaction with ·Cl. All favorably formed indole-radicals further react with O_2 to produce peroxy radicals, which mainly react with NO and HO_2· to form organonitrates, alkoxy radicals and hydroperoxide products. Therefore, the oxidation mechanism of indole is distinct from that of previously reported amines, which primarily form highly oxidized multifunctional compounds, imines or carcinogenic nitrosamines. In addition, the peroxy radicals from the ·OH reaction can form N-(2-formylphenyl)formamide ($C_8H_7NO_2$), for the first time providing evidence for the chemical identity of the $C_8H_7NO_2$ mass peak observed in the ·OH + indole experiments. More importantly, this study is the first to demonstrate that despite forming radicals by abstracting an H-atom at the N-site, carcinogenic nitrosamines were not produced in the indole oxidation reaction.

1 Introduction

Volatile organic compounds (VOCs) play a central role in air quality and climate change as their transformations are highly relevant for the formation of secondary organic aerosols (SOA), toxic air pollutants and ozone (O_3) (Ehn et al., 2014; Karl et al., 2018; Lewis Alastair, 2018; Li et al., 2019; Khare and Gentner, 2018; Ji et al., 2018). Therefore, an accurate description of the atmospheric transformation mechanism and kinetics of VOCs is essential to fully explore the global impacts of VOCs. Despite massive effort to understand the atmospheric fate of VOCs, current mechanism-based atmospheric models often underestimate SOA and O_3 formation quantity. Therefore, the emission inventories or reaction mechanism employed in

the models are either missing some vital primary VOCs or there remain unrevealed reaction mechanism of currently known VOCs. Hence, it is crucial to identify unaccounted reaction pathways of known VOCs or transformation mechanism of unconsidered VOCs with high concentrations.

Organic nitrogen compounds (ONCs) are a subgroup of VOCs that are widely observed in the atmosphere (Silva et al., 2008). Until now, about 160 ONCs have been detected in the atmosphere, accounting for 10% of total gas phase nitrogen (excluding N_2) (Ge et al., 2011; Silva et al., 2008). Due to the adverse effects of ONCs on air quality (formation of particles via acid-base reactions or generation of toxic nitrosamines, nitramines, isocyanic acid and low volatile products via gas phase oxidation), the chemistry of ONCs has gained significant attention in the recent years (Almeida et al., 2013; Chen et al., 2017; Lin et al., 2019; Nielsen et al., 2012; Zhang et al., 2015; Xie et al., 2014; Xie et al., 2015; Xie et al., 2017; Ma et al., 2018a; Ma et al., 2021a; Ma et al., 2019; Shen et al., 2019; Shen et al., 2020). Detailed transformation pathways of a series of ONCs including low-molecular-weight alkyl amines (Nicovich et al., 2015; Xie et al., 2014; Xie et al., 2015; Ma et al., 2021a), aromatic aniline (Xie et al., 2017; Shiels et al., 2021), heterocyclic amines (Sengupta et al., 2010; Ma et al., 2018a; Borduas et al., 2016; Ren and Da Silva, 2019) and amides (Xie et al., 2017; Borduas et al., 2016; Borduas et al., 2015; Bunkan et al., 2016; Bunkan et al., 2015) have been investigated. These studies have shown that the functional groups connected to the NH_x (x = 0, 1, 2) group highly affect the reactivity of ONCs and eventually lead to their different atmospheric impacts. Therefore, the comprehensive understanding the reaction mechanism of ONCs with various functional groups linked to the NH_x group is of great significance to assess the atmospheric impact of ONCs.

Indole is a polyfunctional heterocyclic secondary amine (Laskin et al., 2009). Atmospheric indole has various natural and anthropogenic sources including vegetation, biomass burning, animal husbandry, coal mining, petroleum processing and tobacco industry (Ma et al., 2021b; Cardoza et al., 2003; Yuan et al., 2017; Zito et al., 2015). The global emission of indole is around 0.1 Tg yr^{-1} (Misztal et al., 2015), which is almost equivalent to that of trimethylamine (~ 0.17 Tg yr^{-1}) (Schade and Crutzen, 1995; Yu and Luo, 2014) which has the highest emission among the identified atmospheric amines. A field measurement study found that the concentration of indole can reach 1-3 ppb in ambient air during a springtime flowering event (Gentner et al., 2014). From a structural point of view, the -NH- group of indole is located at 9-center-10-electron delocalized π bonds, possibly altering its reactivity compared to that of previously well-studied aliphatic amines and aniline. Therefore, considering the large atmospheric emission of indole and its unique structure compared to previously studied amines, the reaction mechanism of indole needs to be further evaluated to assess its atmospheric impacts. Furthermore, elucidating the reaction mechanism of indole will add to the fundamental understanding of the transformation mechanism of ONCs.

Hydroxyl radicals (·OH) are considered to be the most important atmospheric oxidants governing the fate of most organic compounds (Macleod et al., 2007). Previous experimental studies have investigated the reaction kinetics and identified the products of the ·OH + indole reaction. Atkinson et al. found that the rate constant (k_{OH}) of the ·OH + indole reaction is 1.54 × 10^{-10} cm^3 $molecule^{-1}$ s^{-1} at 298 K, translating to a 20 min lifetime of indole (Atkinson et al., 1995). Montoya-Aguilera et al. found that isatin and isatoic anhydride are the two dominate monomeric products for ·OH initiated reaction of indole. More importantly, they found that the majority of indole oxidation products can contribute to SOA formation with an effective SOA

yield of 1.3 ± 0.3 under the indole concentration (200 ppb) employed in their chamber study (Montoya-Aguilera et al., 2017). Although the chemical formulas of some of the indole oxidation products have been detected, detailed mechanistic information such as the products branching ratio of the ·OH initiated reaction of indole remains unknown. Additionally, the lack of commercially available standards of some products presents a significant obstacle to identify the exact chemical identity of the products. Therefore, to fully understand the role of indole in SOA formation, it is essential to investigate the detailed atmospheric transformation of indole initiated by ·OH.

Besides reactions with ·OH, reactions with chlorine radicals (·Cl) have been proposed to be an important removal pathway for ONCs due to the identification of new ·Cl continental sources and the high reactivity of ·Cl (Wang et al., 2022; Li et al., 2021; Jahn et al., 2021; Xia et al., 2020; Young et al., 2014; Faxon and Allen, 2013; Riedel et al., 2012; Atkinson et al., 1989; Ji et al., 2013; Thornton et al., 2010; Le Breton et al., 2018). ·Cl initiated atmospheric oxidation of ONCs can lead to the formation of N-centered radicals, once a strong 2-center-3-electron (2c-3e) bond complex has been formed between ·Cl and NH_x (x = 1, 2) (Mckee et al., 1996; Xie et al., 2015; Xie et al., 2017; Ma et al., 2018a). The formed N-centered radicals can further react with NO to form carcinogenic nitrosamines, increasing the atmospheric impact of ONC emissions (Xie et al., 2014; Xie et al., 2015; Xie et al., 2017; Ma et al., 2018a; Ma et al., 2021a; Onel et al., 2014a; Onel et al., 2014b; Nielsen et al., 2012; Da Silva, 2013). As a secondary amine, indole reaction with ·Cl has the possibility of forming N-centered radicals and subsequently forming nitrosamines via the reaction with NO. Since the -NH- group of indole is embedded in a unique chemical environment compared to previously well-studied ONCs, the reaction mechanism of ·Cl with indole remain elusive. In addition, there are only a few studies concerning the reactions of polyfunctional heterocyclic ONCs with ·Cl.

In this work, we investigated the reaction mechanism and kinetics of indole initiated by ·OH and ·Cl by employing a combination of quantum chemical calculations and kinetic modeling. The initial reactions of ·OH/·Cl + indole and the subsequent reactions with O_2 of resulting intermediates were further investigated.

2 Computational Details

2.1 Ab Initio Electronic Structure Calculations

All the geometry optimizations and harmonic vibrational frequency calculations were performed at the M06-2X/6-31+G(d,p) level of theory (Zhao and Truhlar, 2008). Intrinsic reaction coordinate calculations were performed to confirm the connections of each transition state between the corresponding reactants and products. Single point energy calculations were performed at the CBS-QB3 method based on the geometries at the M06-2X/6-31+G(d,p) level of theory (Montgomery et al., 1999). The combination of the M06-2X functional and CBS-QB3 method has successfully been applied to predict radical-molecule reactions (Guo et al., 2020; Ma et al., 2021b; Wang et al., 2018; Wang and Wang, 2016; Wu et al., 2015; Wang et al., 2017; Fu et al., 2020). T_1 diagnostic (Table S2) values in the CCSD(T) calculations within the CBS-QB3 scheme for the intermediates and transition states involved in the key reaction pathways were checked for multireference character. The T_1 diagnostic values for all checked important species in this work are lower than the threshold value of 0.045, indicating the

reliability of applied single reference methods (Rienstra-Kiracofe et al., 2000). In addition, similar to our previous studies, a literature value of 0.8 kcal mol^{-1} for the isolated ·Cl was used to account for the effect of spin-orbit coupling in the ·Cl + indole reaction (Nicovich et al., 2015; Xie et al., 2017; Ma et al., 2018a). Atomic charges of indole and pre-reactive complexes in the ·Cl + indole reaction are obtained by natural bond orbital (NBO) calculations (Reed et al., 1985). All calculations were performed within the Gaussian 09 package (Frisch et al., 2009). Throughout the paper, the symbols "R, RC, PC, TS, IM and P" stand for reactants, pre-reactive complexes, post-reactive complexes, transition states, intermediates and products involved in the reactions, respectively, and their subscripts denote different species. In addition, "A//B" was used to present the computational method, where "A" is the theoretical level for single point energy calculations and "B" is that for geometry optimizations and harmonic frequency calculations.

2.2 Kinetics Calculations

MultiWell-2014.1 and MESMER 5.0 software were employed to investigate the kinetics for short time and long time reaction, respectively (Barker and Ortiz, 2001; Barker, 2001; Glowacki et al., 2012). For the initial reactions of ·OH/·Cl + indole, the reaction kinetics were calculated within the MultiWell-2014.1 program. For the subsequent reactions of resulting primary intermediates, MESMER 5.0 were selected for simulating the reaction kinetics, since it has good performance for long time runs, especially for simulating the variation of the different intermediates over time. Both MultiWell and MESMER codes employs Rice-Ramsperger-Kassel-Marcus (RRKM) theory to calculate the reaction kinetics for reactions with intrinsic energy barriers (Holbrook, 1996; Robinson, 1972). Long-range transition-state theory (LRTST) with a dispersion force potential within the MultiWell-2014.1 program (Barker and Ortiz, 2001) or Inverse Laplace Transformation (ILT) method within the MESMER 5.0 program was employed to calculate the reaction rate constants for the barrierless recombination reactions (from R to RC and P to PC) (Rienstra-Kiracofe et al., 2000). Computational details for performing LRTST and ILT calculation were described in our previous studies (Ma et al., 2021a; Ma et al., 2021b; Guo et al., 2020; Ding et al., 2020b). The parameters used in the LRTST calculations and Lennard-Jones parameters of intermediates estimated by the empirical method proposed by Gilbert and Smith (Gilbert, 1990) are listed in Table S3 and Table S4, respectively. N_2 was selected as the buffer gas, and an average transfer energy of ΔE_d = 200 cm^{-1} was used to simulate the collision energy transfer between active intermediates and N_2. In addition, ΔE_d between 50 - 250 cm^{-1} were selected to study energy transfer parameters effects. For the reactions involving H-abstraction or H-shift, tunneling effects were taken into account in all of the reaction rate constants calculations by using a one-dimensional unsymmetrical Eckart barrier (Eckart, 1930), and were discussed in the Supporting Information (SI). The kinetic calculations were primarily performed at 298 K and 1 atm, with additional ones at 0.1, 0.4 and 0.7 atm in the troposphere relevant range to explore pressure effects. Variation of the energy transfer parameters and pressure resulted in only minor changes (< 0.1%) in the calculated rate coefficients and branching ratios of main reaction pathways (see details in the SI).

3 Results and Discussion

3.1 Initial Reactions of Indole

In principle, ·OH and ·Cl could add to the unsaturated C=C bonds and benzene ring or directly abstract H-atoms connected to either to a C-atom or the N-atom of indole. Considering the planar C_s structure of indole, ·OH and ·Cl addition to one side of indole was only considered here. However, although numerous attempts have been made, we failed to locate the TSs and addition IMs of ·Cl addition to the C2, C3, C4, C7, C8 and C9 sites of indole (the numbering of the atoms is given in Figure 1), suggesting that such additions are in fact unfeasible. Therefore, 7 H-abstraction pathways of ·OH and ·Cl, respectively, 8 ·OH-addition pathways and 2 ·Cl-addition pathways were considered for the ·OH/·Cl + indole reactions. The schematic zero-point energy (ZPE) corrected potential energy surfaces of ·OH/·Cl + indole reactions are presented in Figure 1.

As can be seen from Figure 1, each H-abstraction reaction pathway proceeds through a RC and PC, and the addition pathways through a RC for the ·OH/·Cl + indole reactions. For the H-abstraction pathways, the activation energy (E_a) for the -NH- group for both reactions are at least 2.0 kcal mol^{-1} lower than the corresponding E_a values for the -CH- groups. This indicates that H-abstraction from the -NH- group forming C_8H_6N radicals and H_2O/HCl is the most favorable among all the H-abstraction pathways. In addition, the activation energy for the H-abstraction from the -NH- group in the ·Cl + indole reaction is much lower than the corresponding ·OH + indole reaction. This is consistent with previously reported reactions of other amines with ·OH and ·Cl (Ma et al., 2018a; Ma et al., 2021a; Xie et al., 2014; Xie et al., 2015; Xie et al., 2017; Tan et al., 2021; Ren and Da Silva, 2019; Borduas et al., 2015).

For the addition reactions, the most favorable reaction site differs for the indole + ·OH and indole + ·Cl reactions. Among all 8 ·OH addition pathways, ·OH addition to the C7 site of the C=C bond via $TS_{1\text{-}7}$ forming $IM_{1\text{-}7}$ is the most favorable pathway. Different from the reaction with ·OH, the additions of ·Cl to the C5 and C6 sites to form $IM_{2\text{-}5}$ and $IM_{2\text{-}6}$, respectively, are significantly more favorable. By comparing the E_a values of the addition and H-abstraction pathways for both ·OH/·Cl + indole reactions, it can be concluded that ·OH addition to the C7 site is the most favorable for the ·OH + indole reaction. All the ·OH + indole hydrogen abstraction reactions have high energy barriers. However, the additions of ·Cl to the C5 and C6 sites as well as the -NH- H-abstraction are all favorable due to their very lower E_a values for the ·Cl + indole reaction.

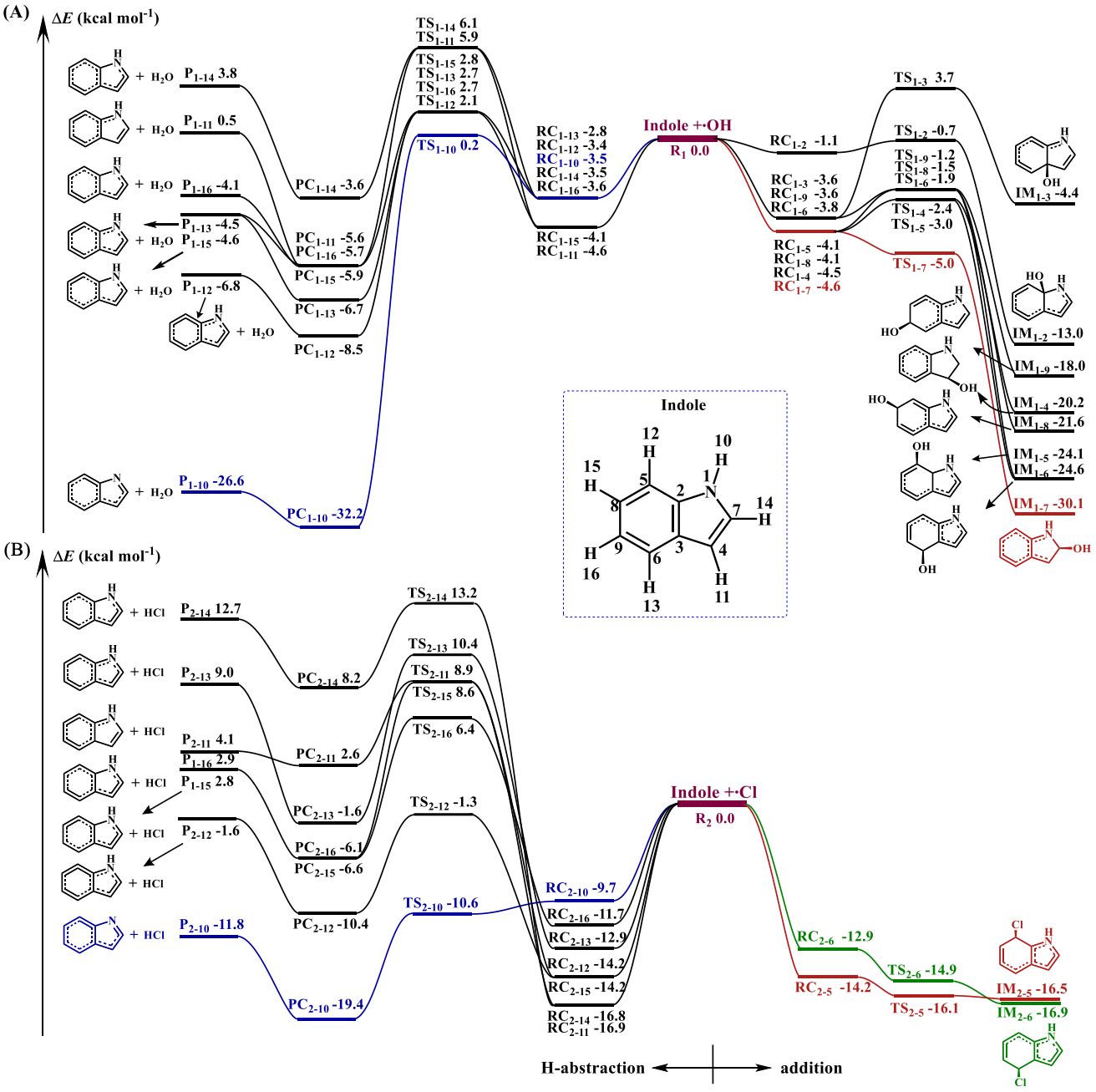


Figure 1: Schematic ZPE-corrected potential energy surface for the reactions of indole + ·OH (A) and indole + ·Cl (B) at the CBS-QB3//M062X/6-31+g(d,p) level of theory. The total energy of the reactants indole + ·OH/·Cl are set to zero, respectively (reference state).

Interestingly, we found that all the pathways for the indole + ·Cl reaction can proceed via a stable 2c-3e bonded RC, which is different from that of the ·OH + indole reaction. Among all 2c-3e bonded RCs, only $RC_{2\text{-}10}$ from the -NH- abstraction pathway is formed between the N-atom and ·Cl, while the others are formed between the C-atom and ·Cl. Note that $RC_{2\text{-}11}$,

which forms from C-atom and ·Cl, is the most stable among all the formed RCs in the ·Cl + indole reaction. To the best of our knowledge, this is the first time that such a stable 2c-3e bonded RC has been identified between the C-atom and ·Cl. In addition, the energy of $RC_{2\text{-}10}$ is higher than that of the traditional 2c-3e bonded RCs formed from alkylamine and ·Cl, which would result from the delocalization of lone pair electrons of the N-atom. By analyzing the NBO charges of these nine RCs (Table S5), we found that significant charge transfer occurs between ·Cl and indole. The charge at Cl atom for $RC_{2\text{-}5}$, $RC_{2\text{-}6}$, $RC_{2\text{-}10}$, $RC_{2\text{-}11}$, $RC_{2\text{-}12}$, $RC_{2\text{-}13}$, $RC_{2\text{-}14}$, $RC_{2\text{-}15}$ and $RC_{2\text{-}16}$ are -0.35 *e*, -0.33 *e*, -0.31 *e*, -0.39 *e*, -0.35 *e*, -0.33 *e*, -0.39 *e*, -0.35 *e* and -0.33*e*, respectively, indicating that all RCs are charge-transfer complexes. Similar charge-transfer complexes were also found in our previous study of the ·Cl + piperazine reaction (Ma et al., 2018a).

With the master equation theory, the overall rate constants (k_{OH} and k_{Cl}) and branching ratios (Γ) for all H-abstraction and ·OH/·Cl-addition pathways involved in the ·OH/·Cl + indole reactions were calculated at 298 K and 1 atm. The calculated k_{OH} and k_{Cl} values of indole are 7.9×10^{-11} cm^3 $molecule^{-1}$ s^{-1} and 2.9×10^{-10} cm^3 $molecule^{-1}$ s^{-1}, respectively. The calculated k_{OH} value is close to the available experimental value of 1.5×10^{-10} cm^3 $molecule^{-1}$ s^{-1} (Atkinson et al., 1995), supporting the reliability of employed computational methods. Over the temperature range 230-330 K (Ma et al., 2018b), the calculated k_{OH} and k_{Cl} values have a negative correlation with temperature (Figure S1). Based on the calculated Γ values of the ·OH/·Cl + indole reactions (Table 1), it can be concluded that $IM_{1\text{-}7}$ (77%) is the main product for ·OH + indole reaction, and $IM_{2\text{-}5}$ (31%), $IM_{2\text{-}6}$ (46%) and $P_{2\text{-}10}$ (C_8H_6N radicals + HCl) (23%) are the main products for ·Cl + indole reaction. In addition, the calculated Γ values of $IM_{1\text{-}7}$, $IM_{2\text{-}5}$, $IM_{2\text{-}6}$ and $P_{2\text{-}10}$ (C_8H_6N radicals + HCl) change negligibly with the variation of temperature, pressure and energy transfer parameters (SI). Therefore, we mainly considered the further transformation of $IM_{1\text{-}7}$, $IM_{2\text{-}5}$, $IM_{2\text{-}6}$ and C_8H_6N radicals in the following part.

Table 1. Calculated branching ratios (Γ) for the indole + ·OH/·Cl reactions at 298 K and 1 atm.

Pathways	Species	Γ	Species	Γ	Species	Γ
·OH + Indole	$IM_{1\text{-}2}$	0	$IM_{1\text{-}3}$	0	$IM_{1\text{-}4}$	5%
	$IM_{1\text{-}5}$	12%	$IM_{1\text{-}6}$	3%	$IM_{1\text{-}7}$	77%
	$IM_{1\text{-}8}$	1%	$IM_{1\text{-}9}$	1%	$P_{1\text{-}10}$	1%
	$P_{1\text{-}11}$	0	$P_{1\text{-}12}$	0	$P_{1\text{-}13}$	0
	$P_{1\text{-}14}$	0	$P_{1\text{-}15}$	0	$P_{1\text{-}16}$	0
·Cl + Indole	$IM_{2\text{-}5}$	31%	$IM_{2\text{-}6}$	46%	$P_{2\text{-}10}$	23%
	$P_{2\text{-}11}$	0	$P_{2\text{-}12}$	0	$P_{2\text{-}13}$	0
	$P_{2\text{-}14}$	0	$P_{2\text{-}15}$	0	$P_{2\text{-}16}$	0

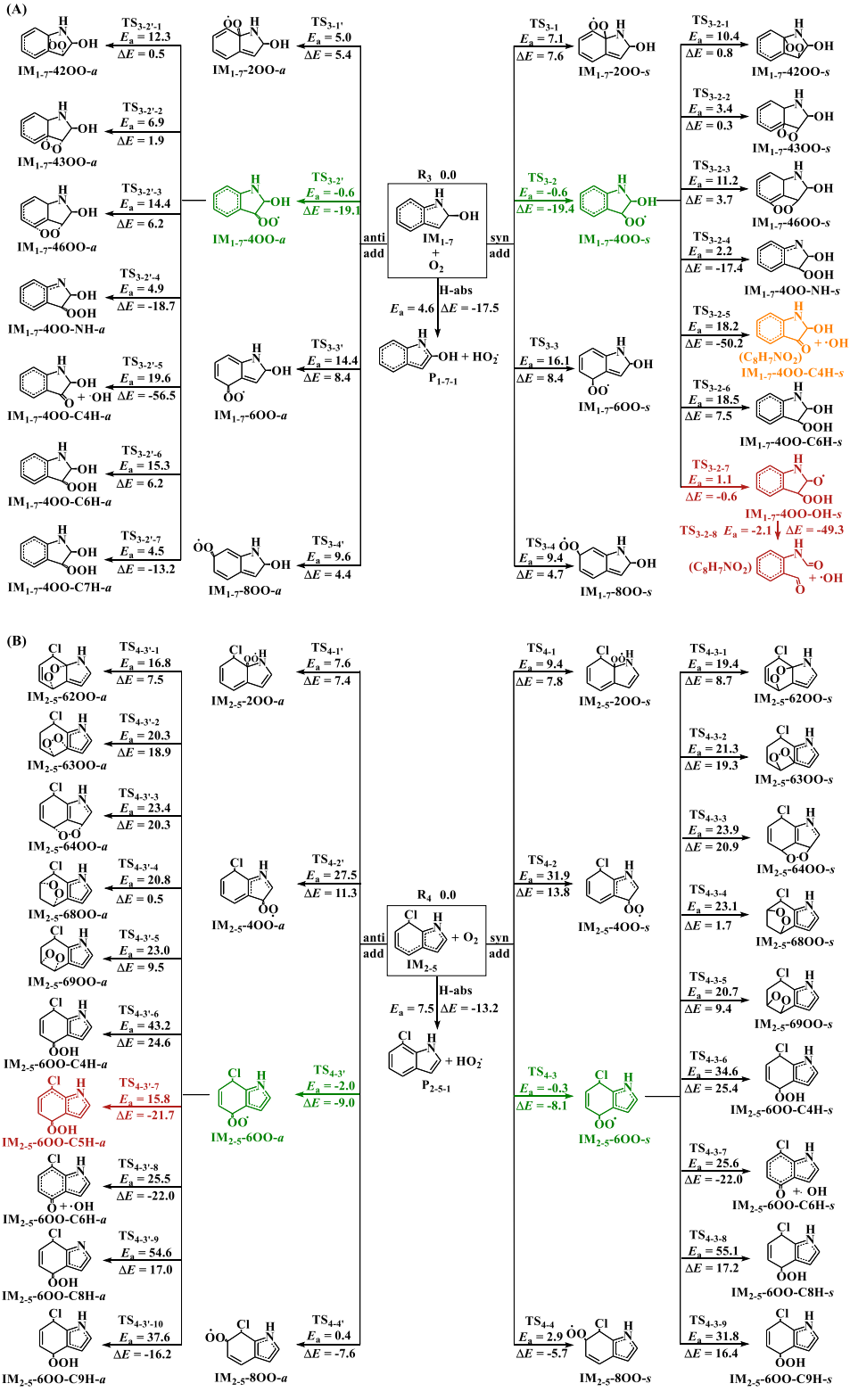
3.2 Subsequent Reactions of Addition Intermediates

Similar to other C-centered radicals (Zhang et al., 2012; Guo et al., 2020; Ma et al., 2021b; Yu et al., 2016; Yu et al., 2017; Ji et al., 2017; Ding et al., 2020a), the intermediates $IM_{1\text{-}7}$, $IM_{2\text{-}5}$ and $IM_{2\text{-}6}$ will subsequently react with O_2. Two different

pathways (Figure 2) were considered for the reactions of the intermediates IM_{1-7}, IM_{2-5} and IM_{2-6} with O_2. One is the direct hydrogen abstraction by O_2 from the C site connecting to the -OH or -Cl group forming P_{1-7-1} ($C_8H_7NO + HO_2\cdot$), P_{2-5-1} ($C_8H_6NCl + HO_2\cdot$) and P_{2-6-1} ($C_8H_6NCl + HO_2\cdot$). The other is the O_2 addition to the C sites with high spin density (see spin density distribution in Table S10) of the intermediates IM_{1-7}, IM_{2-5} and IM_{2-6} to form peroxy radicals *Q-i*OO-*a/s*, where *Q* stands for intermediates IM_{1-7}, IM_{2-5} and IM_{2-6}, *i* stands for the numbering of the C-positions where O_2 is added. The O_2 molecule can be added to the same (-*syn*, abbreviated as -*s*) and opposite (-*anti*, abbreviated as -*a*) sides of the plane relative to -OH or -Cl group. The C2, C4, C6 and C8 sites of IM_{1-7}, C2, C4, C6 and C8 sites of IM_{2-5} and C3, C5, C7 and C9 sites of IM_{2-6} are high spin density sites susceptible for O_2 addition.

As can be seen from the energetic data shown in Figure 2, O_2 addition to the C4 site of IM_{1-7} to form IM_{1-7}-4OO-*a/s* (-0.6/-0.6 kcal mol^{-1}), C6 site of IM_{2-5} to form IM_{2-5}-6OO-*a/s* (-0.3/-2.0 kcal mol^{-1}) and C5 site of IM_{2-6} to form IM_{2-6}-5OO-*a/s* (2.0/1.7 kcal mol^{-1}) are the most favorable among all possible entrance pathways for the respective reactions. It deserves mentioning that the formation energy (ΔE) of IM_{2-5}-6OO-*a/s* and IM_{2-6}-5OO-*a/s* are only about 9.0 kcal mol^{-1}, which could indicate that they likely re-dissociate back to the reactants IM_{2-5}/IM_{2-6} and O_2, if IM_{2-5}-6OO-*a/s* and IM_{2-6}-5OO-*a/s* does not rapidly transform to other species.

For the further transformation of the formed peroxy radicals IM_{1-7}-4OO(-*a/s*), IM_{2-5}-6OO(-*a/s*) and IM_{2-6}-5OO(-*a/s*), two transformation pathways were identified. The first is cyclization reactions where the terminal O-atom of -OO attacks the different C-positions to form bicycle radicals *Q-ij*OO(-*a/s*) (*j* stands the number of the C-positions attacked by terminal O-atom). The second is H-shifts from -OH, -NH- and different -CH- sites to the terminal O-atom to form various hydroperoxide radicals *Q-i*OO-OH(-*a/s*), *Q-i*OO-NH(-*a/s*) and *Q-i*OO-C*k*H(-*a/s*) (*k* stands the number of the C-positions from which H is shifted), respectively. For IM_{1-7}-4OO(-*a/s*) and IM_{2-5}-6OO(-*a/s*), forming IM_{1-7}-4OO-OH-*s* and IM_{2-5}-6OO-C5H-*a* via H-shift reactions are the most favorable, respectively. However, for IM_{2-6}-5OO(-*a/s*), the cyclization reaction forming IM_{2-6}-52OO-*a* is the most favorable. It is noted that the formed IM_{1-7}-4OO-OH-*s* from IM_{1-7}-4OO(-*a/s*) can barrierlessly transform to form $C_8H_7NO_2$ (N-(2-formylphenyl)formamide) and ·OH (collectively denoted $P_{1-7-4-1}$) via concerted C-C and O-O bonds rupture. The further transformation of the peroxy radicals IM_{1-7}-4OO(-*a/s*), IM_{2-5}-6OO(-*a/s*) and IM_{2-6}-5OO(-*a/s*) need to overcome barriers above 20.5 kcal mol^{-1} (relative to their respective peroxy radicals), indicating that the further transformation of IM_{1-7}-4OO(-*a/s*), IM_{2-5}-6OO(-*a/s*) and IM_{2-6}-5OO(-*a/s*) should be very slow.


(A)
IM1-7-42OO-a
IM1-7-43OO-a
IM1-7-46OO-a
IM1-7-4OO-NH-a
IM1-7-4OO-C4H-a
IM1-7-4OO-C6H-a
IM1-7-4OO-C7H-a
IM1-7-2OO-a
IM1-7-4OO-a
IM1-7-6OO-a
IM1-7-8OO-a
R3 0.0
IM1-7 + O2
anti add
syn add
H-abs
Ea = 4.6 ΔE = -17.5
P1-7-1
IM1-7-2OO-s
IM1-7-4OO-s
IM1-7-6OO-s
IM1-7-8OO-s
IM1-7-42OO-s
IM1-7-43OO-s
IM1-7-46OO-s
IM1-7-4OO-NH-s
IM1-7-4OO-C4H-s
IM1-7-4OO-C6H-s
IM1-7-4OO-OH-s
(B)
IM2-5-62OO-a
IM2-5-63OO-a
IM2-5-64OO-a
IM2-5-68OO-a
IM2-5-69OO-a
IM2-5-6OO-C4H-a
IM2-5-6OO-C5H-a
IM2-5-6OO-C6H-a
IM2-5-6OO-C8H-a
IM2-5-6OO-C9H-a
IM2-5-2OO-a
IM2-5-4OO-a
IM2-5-6OO-a
IM2-5-8OO-a
R4 0.0
IM2-5 + O2
anti add
syn add
H-abs
Ea = 7.5 ΔE = -13.2
P2-5-1
IM2-5-2OO-s
IM2-5-4OO-s
IM2-5-6OO-s
IM2-5-8OO-s
IM2-5-62OO-s
IM2-5-63OO-s
IM2-5-64OO-s
IM2-5-68OO-s
IM2-5-69OO-s
IM2-5-6OO-C4H-s
IM2-5-6OO-C6H-s
IM2-5-6OO-C8H-s
IM2-5-6OO-C9H-s

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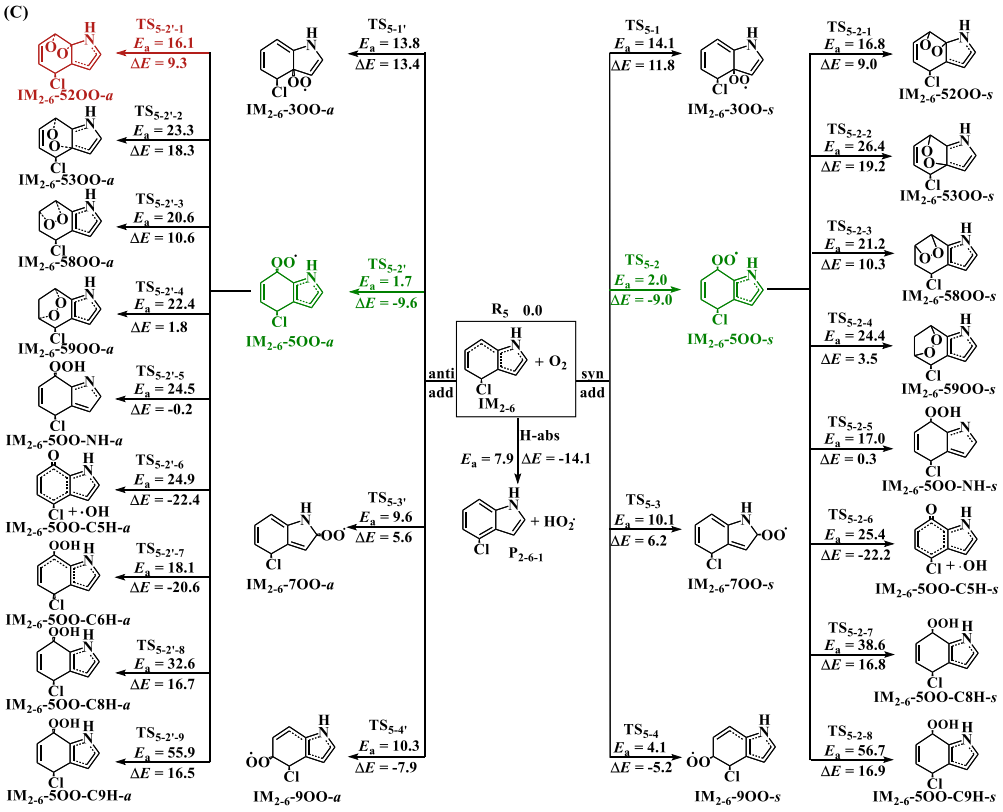

Figure 2: Reaction pathways and corresponding energetic data for the reactions of IM_{1-7} (A), IM_{2-5} (B) and IM_{2-6} (C) with O_2. Units are in kcal mol^{-1}.

Based on the energetic data of the favorable reaction pathways, MESMER modeling was employed to investigate the reaction rate constants and fractional yields for the reactions of IM_{1-7}, IM_{2-5}, and IM_{2-6} with O_2. Similar to previous studies (Guo et al., 2020; Ma et al., 2021a; Ma et al., 2021b; Zhang et al., 2012; Fu et al., 2020), bimolecular reactions with NO/HO_2· are considered as competitive pathways for the unimolecular reactions of the peroxy radicals IM_{1-7}-4OO(-*a*/*s*), IM_{2-5}-6OO(-*a*/*s*) and IM_{2-6}-5OO(-*a*/*s*) by simply adding their pseudo-first-order rate constants into the master equation modeling. Here, applied pseudo first order rate constants for peroxy radicals (IM_{1-7}-4OO(-*a*/*s*), IM_{2-5}-6OO(-*a*/*s*) and IM_{2-6}-5OO(-*a*/*s*)) reaction with NO and HO_2· are 0.06 s^{-1} and 0.02 s^{-1}, respectively, corresponding to reactions occurring at 200 ppt NO and 50 ppt HO_2· conditions (Hofzumahaus et al., 2009; Yu et al., 2020; Praske et al., 2018). The reactions of peroxy radicals with NO and HO_2· should form organonitrate/alkoxy radicals (collectively denoted NO-P_n, where *n* marks products from the different peroxy radical reactions) and hydroperoxide (HO_2-P_n), respectively. Pseudo-first-order rate constants of IM_{1-7}, IM_{2-5}, and IM_{2-6} with O_2 are calculated to be 3.0×10^7 s^{-1}, based on the reaction rate constants of IM_{1-7}, IM_{2-5}, and IM_{2-6} with O_2 (6.0×10^{-12} cm^3 $molecule^{-1}$ s^{-1}) and the concentration of O_2 ($[O_2] = 5.0 \times 10^{18}$ molecule cm^{-1}). The simulated time-dependent fractional yields are presented in Figure 3.

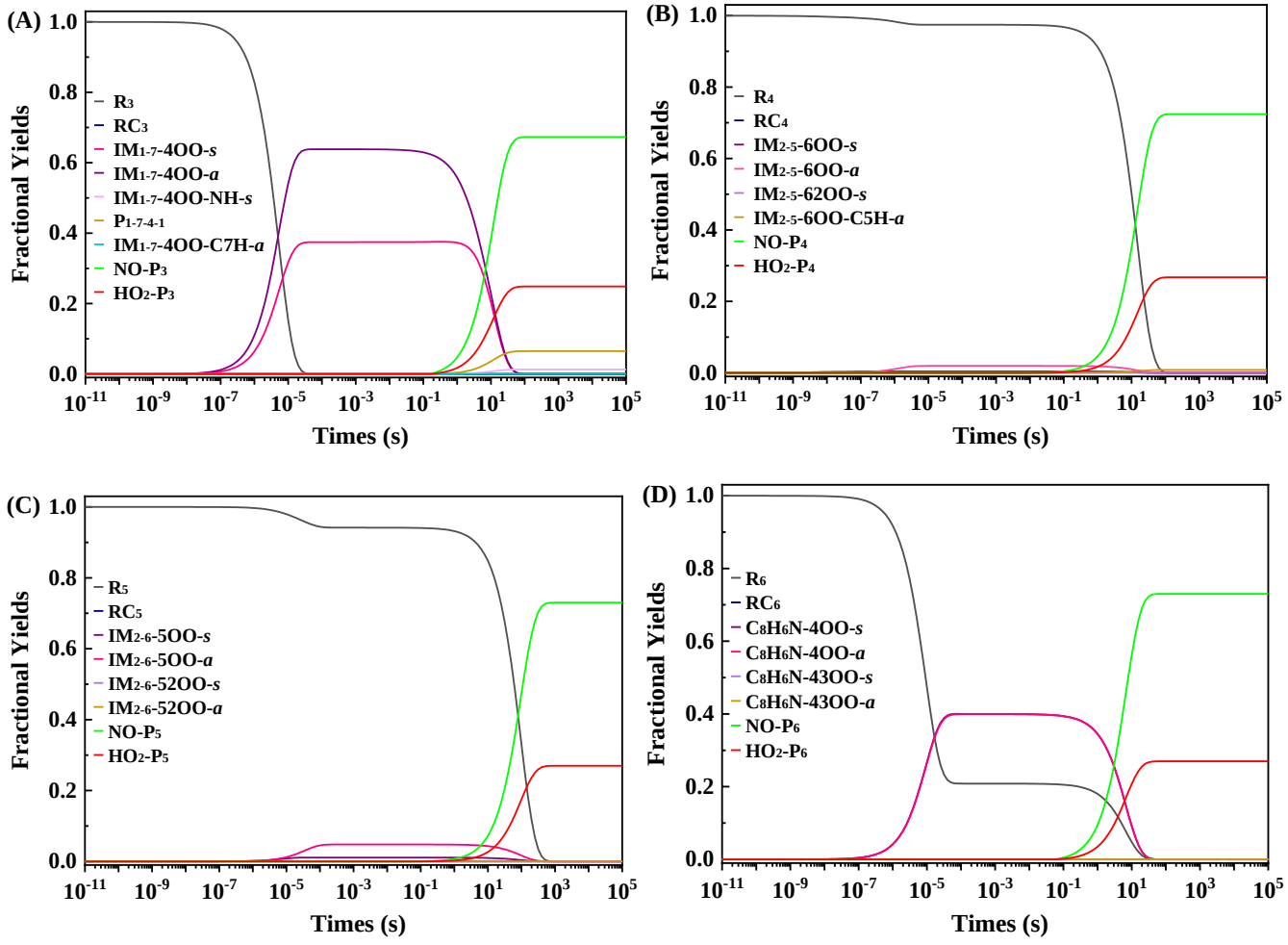

Figure 3: Calculated fractional yields of species (at 200 ppt NO and 50 ppt HO_2· conditions) as a function of time in the reactions of IM_{1-7} (A), IM_{2-5} (B), IM_{2-6} (C) and C_8H_6N (D) with O_2 at 298 K and 1 atm.

As can be seen in Figure 3, after 100 s, the reactions of IM_{1-7}, IM_{2-5} and IM_{2-6} with O_2 mainly form the organonitrate/alkoxy radicals NO-P_3 ($C_8H_8N_2O_3$/$C_8H_8NO_2$·), NO-P_4 ($C_8H_7N_2O_3Cl$/C_8H_7NClO·) and NO-P_5 ($C_8H_7N_2O_3Cl$/C_8H_7NClO·), followed by the formation of hydroperoxide HO_2-P_3 ($C_8H_9NO_3$), HO_2-P_4 ($C_8H_8NO_2Cl$) and HO_2-P_5 ($C_8H_8NO_2Cl$), respectively. For the reactions of IM_{2-5} and IM_{2-6} with O_2, the main products formed are NO-$P_{4/5}$ and HO_2-$P_{4/5}$. In contrast, the IM_{1-7} + O_2 reaction also lead to the fragmental products $P_{1-7-4-1}$ ($C_8H_7NO_2$ and ·OH) besides the main products NO-P_3 and HO_2-P_3. This difference in product branching ratios results from the lower unimolecular reaction energy barrier of the peroxy radicals IM_{1-7}-4OO(-*a*/*s*) from the reaction of IM_{1-7} with O_2 than those of IM_{2-5}-6OO(-*a*/*s*) and IM_{2-6}-5OO(-*a*/*s*) from the reactions of IM_{2-5} and IM_{2-6} with O_2. It should be noted that the $C_8H_7NO_2$ product has been detected in previous experimental study of the ·OH + indole reaction (Montoya-Aguilera et al., 2017), supporting the validity of our computational results.

An obvious difference for these three reactions is that the reaction of $IM_{1\text{-}7}$ with O_2 can form peroxy radicals $IM_{1\text{-}7}$-4OO(-*a/s*) with high yields during the reactions. However, the yields of the corresponding peroxy radicals $IM_{2\text{-}5}$-6OO(-*a/s*) and $IM_{2\text{-}6}$-5OO(-*a/s*) from the reactions of $IM_{2\text{-}5}$ and $IM_{2\text{-}6}$ with O_2 are low. The difference mainly originates from the difference in the formation energy of these three peroxy radicals as shown in Figure 2. The ΔE values of $IM_{1\text{-}7}$-4OO(-*a/s*)(-19.1/-19.4 kcal mol^{-1}) are much more lower than those of $IM_{2\text{-}5}$-6OO(-*a/s*)(-9.0/-8.1 kcal mol^{-1}) and $IM_{2\text{-}6}$-5OO(-*a/s*)(-9.6/-9.0 kcal mol^{-1}). As discussed above, the high formation energy of $IM_{2\text{-}5}$-6OO(-*a/s*) and $IM_{2\text{-}6}$-5OO(-*a/s*) should make $IM_{2\text{-}5}$-6OO(-*a/s*) and $IM_{2\text{-}6}$-5OO(-*a/s*) return back to the reactants, explaining the reason for the lower yields of $IM_{2\text{-}5}$-6OO(-*a/s*) and $IM_{2\text{-}6}$-5OO(-*a/s*).

3.3 Subsequent Reactions of C_8H_6N radicals from the H-abstraction pathway

Here, the bimolecular reaction with O_2 was mainly considered for C_8H_6N radicals as its sole atmospheric fate. It was found that the spin density distribution was mainly centered at the C atoms (C4 (0.662), C6 (0.261), C8 (0.178)) and N atom (0.256), indicating that the C_8H_6N radical is delocalized. This is in contrast to previously studied N-centered radicals formed from alkylamines oxidation, which are highly localized (Xie et al., 2015; Xie et al., 2014; Ma et al., 2018a; Tan et al., 2021; Borduas et al., 2015). Therefore, O_2 addition to the C4, C6, C8 and N1 sites (including attack from both sides) are considered for the reaction of the C_8H_6N radicals with O_2. As can be seen from Figure 4, O_2 additions to the C4 site of the C_8H_6N radicals forming C_8H_6N-4OO-*a/s* with E_a of -0.3 kcal mol^{-1} are the most favorable, translating to pseudo-first-order reaction rate constants of 3.0×10^7 s^{-1}. Such rate constants are about 7 orders of magnitude higher than that of typical N-centered radicals reacting with NO even under very high NO concentration (5 ppb). Therefore, C_8H_6N radicals does not react with NO to form carcinogenic nitrosamines in any appreciable amount, which is different from the previously reported reaction mechanism of N-centered radicals formed from the reactions of alkylamines with ·Cl (Xie et al., 2015; Xie et al., 2014; Ma et al., 2018a). To the best of our knowledge, this is the first study to reveal that despite forming radicals by abstracting an H-atom at the N-site, carcinogenic nitrosamines were not produced in the indole oxidation reaction.

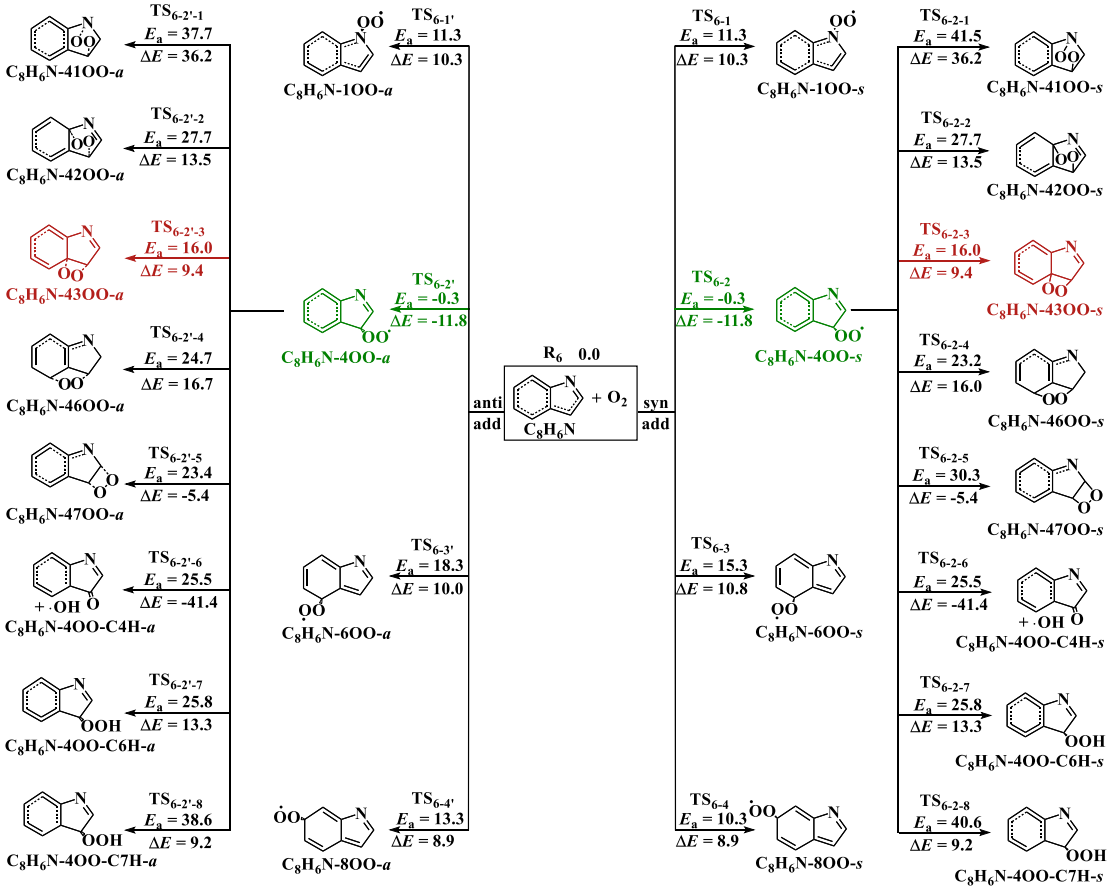

Figure 4: Reaction pathways and corresponding energetic data for the reactions of C_8H_6N radicals with O_2. Units are in kcal mol^{-1}.

For the transformation of the formed C_8H_6N-4OO-*a/s* radicals, the ring closure reaction to form C_8H_6N-43OO-*a/s* is the most favorable, but still needs to overcome a 27.8 kcal mol^{-1} energy barrier, therefore the further transformation of the formed C_8H_6N-4OO-*a/s* should proceed very slowly. The C_8H_6N-4OO-*a/s* should mainly react with NO and HO_2· to form NO-P_6 and HO_2-P_6. Detailed kinetics calculations (Figure 3D) further confirm that the reaction of C_8H_6N radicals with O_2 mainly form NO-P_6 and HO_2-P_6 under 200 ppt NO and 50 ppt HO_2· conditions.

4 Comparison with Available Experimental Results and Atmospheric Implications.

This study found that ·OH/·Cl initiated reactions of indole mainly form organonitrates, alkoxy radicals and hydroperoxide products with N-(2-formylphenyl)formamide ($C_8H_7NO_2$) as a minor product at 200 ppt NO and 50 ppt HO_2· conditions. The formed closed-shell products have high oxygen-to-carbon ratios compared to indole and therefore are expected to have lower vapor pressures, likely being first generation products that can be further oxidized and contribute to the formation of SOA.

With our findings, a comparison was made with the available experimental study on ·OH initiated reaction of indole. The calculated k_{OH} values (7.9×10^{-11} cm^3 $molecule^{-1}$ s^{-1}) of indole is consistent with the experimental value (15×10^{-11} cm^3 $molecule^{-1}$ s^{-1}) (Atkinson et al., 1995), indicating the reliability of applied theoretical methods. A signal with the molecular formula $C_8H_7NO_2$ has been observed in the mass spectrum in an experimental study (Montoya-Aguilera et al., 2017), supporting the formation of the predicted N-(2-formylphenyl)formamide. To the best of our knowledge, this study is the first to reveal that the chemical identity of the mass spectrum signal as N-(2-formylphenyl)formamide, as opposed to the proposed 3-oxy-2-hydroxy-indole. In addition, monomeric products (isatin and isatoic anhydride) and dimer products has not been observed in our computational study. We speculate that they may be produced from the subsequent conversion of the formed alkoxy radicals, multi-generation reactions of organonitrates and hydroperoxide and self/cross reactions of peroxy radicals (RO_2 + RO_2). Therefore, further studies are warranted to investigate the subsequent transformation of the formed alkoxy radicals, organonitrates and hydroperoxide, and the RO_2 + RO_2 reactions, to accurately describe the atmospheric impact of indole.

The calculated k_{Cl} value of the indole + ·Cl reaction is a factor of 3.7 higher than that of the indole + ·OH reaction, and is close to the k_{Cl} values for the reactions of alkylamines, heterocyclic amines and amides with ·Cl (Xie et al., 2017; Xie et al., 2015; Ma et al., 2018a; Nicovich et al., 2015). The contribution of ·Cl to the transformation of indole is calculated to be 3.6-36% that of ·OH, assuming ·Cl concentrations equal to 1-10% of that of ·OH (Wang and Ruiz, 2017; Nicovich et al., 2015; Xie et al., 2017; Xie et al., 2015; Ma et al., 2018a). Therefore, ·Cl plays an important role in the overall transformation of indole. More importantly, ·Cl initiated reaction of indole does not lead to the formation of carcinogenic nitrosamines although ·Cl can favorably abstract the H-atom from N-site to form C_8H_6N radicals, which is a plausible precursor of carcinogenic nitrosamines. Hence, to the best of our knowledge, this is the first study to reveal that despite forming radicals by abstracting an H-atom at the N-site, carcinogenic nitrosamines were not produced in the indole oxidation reaction. This is most likely caused by the delocalized character of the formed C_8H_6N radicals due to the existence of the adjacent unsaturated bonds. Therefore, this study further confirm that the functional groups connected to the NH_x (x = 1, 2) group highly affect the atmospheric fate of ONCs. Further studies should be performed to investigate the structure-activity relationship of ·Cl initiated reactions of ONCs to comprehensively evaluate their atmospheric impacts.

Data availability. The data in this article are available from the corresponding author upon request (maff@dlut.edu.cn, hbxie@dlut.edu.cn).

Author contribution. FFM and HBX designed research; JWX, FFM and HBX performed research; JWX, FFM and HBX analyzed data; JWX, FFM, JE, HBX and JWC wrote the paper; and FFM, HBX, JE and JWC reviewed and revised the paper.

Competing interests. The authors declare that they have no conflict of interest.

Acknowledgements. We thank Dr. Struan H. Robertson (Dassault Systèmes) for the discussion on the MESMER simulations. The study was supported by the LiaoNing Revitalization Talents Program (XLYC1907194), National Natural Science Foundation of China (21876024), the Major International (Regional) Joint Research Project (21661142001) and Supercomputing Center of Dalian University of Technology.

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