



Investigation of the limonene photooxidation by OH at different NO concentrations in the atmospheric simulation chamber SAPHIR

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Abstract. The oxidation of limonene by the hydroxyl (OH) radical and ozone (O₃) was investigated in the atmospheric simulation chamber SAPHIR in experiments performed at different nitric oxide (NO) mixing ratios from nearly zero up to 10
20 ppbv. For the experiments dominated by OH oxidation the formaldehyde (HCHO) yield was experimentally determined and found to be (12±3), (13±3), and (32±5) % for experiments with low (~0.1 ppbv), medium (~0.3 ppbv), and high NO (5 to 10 ppbv), respectively. The yield in an ozonolysis-only experiment was (10±1)%, which agrees with previous laboratory studies. The experimental yield of the first generation organic nitrates from limonene-OH oxidation is calculated as (34±5)%, about 11% higher than the value in the Master Chemical Mechanism (MCM), which is derived from structure-activity-relationships
25 (SAR). Time series of measured radicals, trace-gas concentrations, and OH reactivity are compared to results from zero-dimensional chemical box model calculations applying the MCM v3.3.1. Modelled OH reactivity is 5 to 10 s⁻¹ (25% to 33% of the OH reactivity at the start of the experiment) higher than measured values at the end of the experiments at all chemical conditions investigated, suggesting either that there are unaccounted loss processes of limonene oxidation products or that products are less reactive toward OH. In addition, model calculations underestimate measured hydroperoxyl radical (HO₂)
30 concentrations by 20% to 90% and overestimate organic peroxy radical (RO₂) concentrations by 50% to 300%. Largest deviations are found in low-NO experiments and in the ozonolysis experiment. An OH radical budget analysis, which uses only measured quantities, shows that the budget is closed in most of the experiments. A similar budget analysis for RO₂ radicals



suggests that an additional RO₂ loss rate of about $(1 - 6) \times 10^{-2} \text{ s}^{-1}$ for first-generation RO₂ is required to match the measured RO₂ concentrations in all experiments. Sensitivity model runs indicate that additional reactions converting RO₂ to HO₂ at a rate of about $(1.7 - 3.0) \times 10^{-2} \text{ s}^{-1}$ would improve the model-measurement agreement of NO_x, HO₂, RO₂ concentrations, and OH reactivity. Reaction pathways that could lead to the production of additional OH and HO₂ are discussed, which include isomerisation reactions of RO₂ from the oxidation of limonene, different branching ratios for the reaction of RO₂ with HO₂, and a faster rate constant for RO₂ recombination reactions. As the exact chemical mechanisms of the additional HO₂ and OH sources could not be identified, further work needs to focus on quantifying organic product species and organic peroxy radicals from limonene oxidation.

1 Introduction

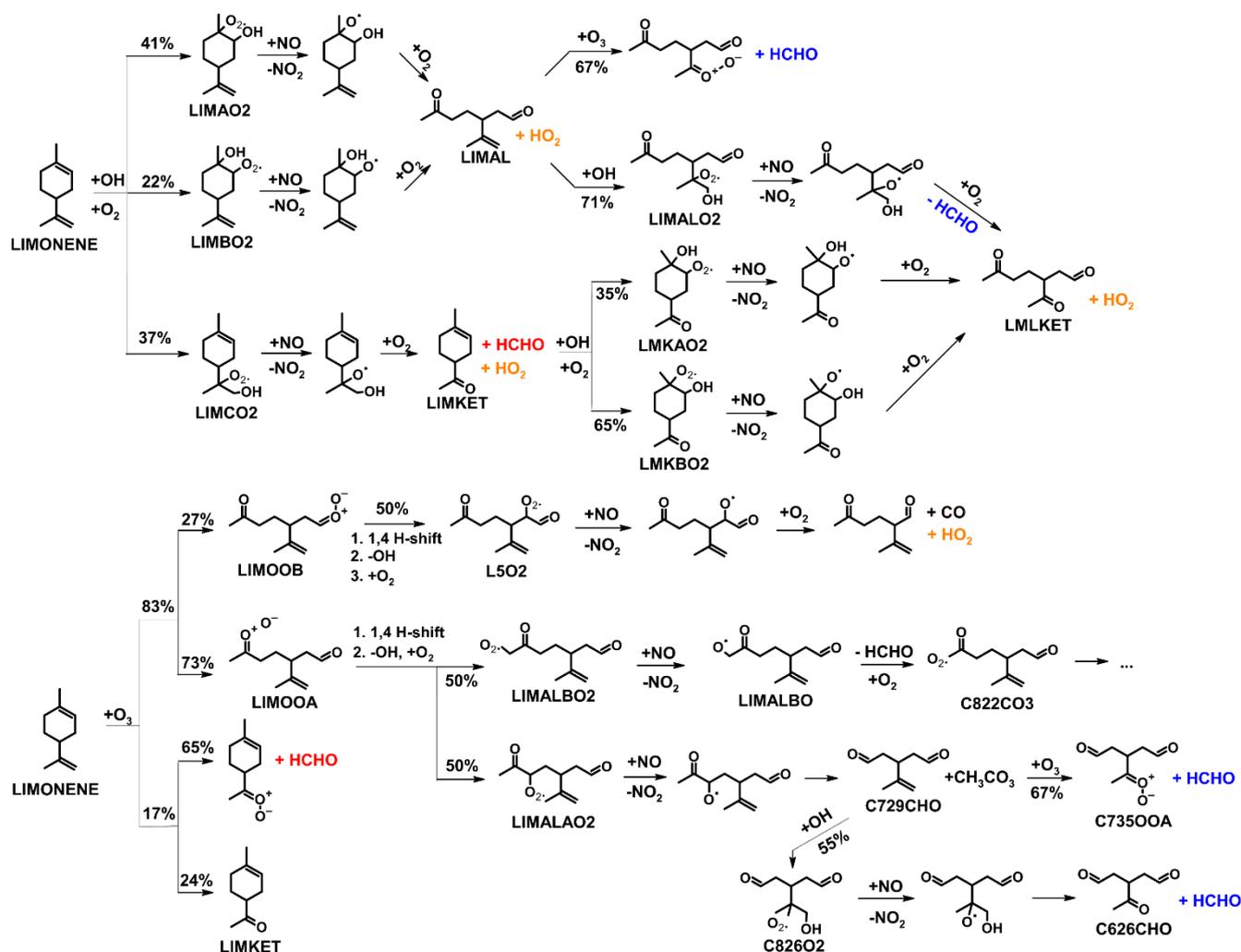
About 1000 Tg of biogenic volatile organic compounds (BVOCs) are emitted into the atmosphere every year. Approximately half of the emissions comprise isoprene while 15% consist of monoterpene species (Guenther et al., 2012). Among monoterpenes, limonene is the fourth most abundant species, which comprises about 10% of the total monoterpene emissions (Sindelarova et al., 2014). Apart from biogenic sources, limonene is also emitted from anthropogenic sources such as air fresheners and cleaning products (Liu et al., 2004; Nazaroff and Weschler, 2004; McDonald et al., 2018; Gkatzelis et al., 2021a, b). Limonene has two carbon-carbon double bonds, which make it reactive towards atmospheric oxidants such as ozone (O₃), hydroxyl radicals (OH), and nitrate radicals (NO₃) (Calogirou et al., 1999; Atkinson and Arey, 2003).

During daytime, the most important oxidant is the hydroxyl radical (OH). The reaction between the olefin group and the OH radical yields a β-hydroxyalkyl radical, which subsequently reacts with an oxygen molecule (O₂) under atmospheric conditions to form an organic peroxy radical (RO₂) (Fig. 1). In the presence of nitric oxide (NO), RO₂ is converted to an alkoxy radical (RO) and thereby NO is oxidized to nitrogen dioxide (NO₂). NO₂ can be photolyzed back to NO together with an oxygen atom, which then reacts with an oxygen molecule producing ozone. This mechanism is the most relevant source for tropospheric ozone. Alkoxy radicals are very reactive and they often quickly decompose forming carbonyl products and a hydroperoxyl radical (HO₂) or further organic radicals (Orlando et al., 2003). OH is then regenerated through the reaction between HO₂ and NO, closing the radical reaction cycle and producing another ozone molecule from the photolysis of nitrogen dioxide. In polluted environments with high NO_x (=NO+NO₂) emissions such as in urban areas (NO concentration > 1000 pptv; e.g., Dusanter et al. (2009) and Lu et al. (2017)), the reaction between RO₂ and NO is often the dominant pathway, through which OH is regenerated. In remote regions with low NO mixing ratios (< 200 pptv; e.g., Ren et al. (2006) and Whalley et al. (2011)), RO₂ predominantly reacts with HO₂ and forms an organic peroxide molecule. This process terminates the radical chain. The organic peroxide produced during this process can potentially add to secondary organic aerosols (SOA) (Surratt et al., 2006). RO₂ radicals can also react with other RO₂ radicals. The reaction rate constant is usually low (*k* ranges from 10⁻¹⁵ to 10⁻¹¹ cm³ s⁻¹; (Tyndall et al. (2001); Jenkin et al. (2019))) and the RO₂ loss rate is often small for atmospheric RO₂ concentrations (RO₂



65 $\sim 10^8 \text{ cm}^{-3}$; Tan et al. (2018)). Products of the RO_2 recombination reactions are either alkoxy radicals or oxidized organic compounds containing alcohol or carbonyl functional groups, which terminate the radical reaction chain.

During the last decade, there has been increasing evidence from laboratory studies as well as theoretical studies that also unimolecular reactions of RO_2 can be of importance in the atmosphere. Unimolecular H-shift reactions of RO_2 can promote the formation of low volatility organic compounds as the number of oxygen atoms in the molecule quickly increases during the process (Jokinen et al., 2014). In some of the unimolecular reactions, HO_x ($=\text{OH}+\text{HO}_2$) radicals can be regenerated, so that
70 this reaction pathway becomes a radical propagation reaction that does not require the presence of NO. This has been, for example, shown for isoprene (Peeters and Müller, 2010; Crouse et al., 2011; Peeters et al., 2014; Berndt et al., 2019; Novelli et al., 2020) and methacrolein (Crouse et al. 2012, Fuchs et al., 2014). The reaction rate constant of H-shift unimolecular reactions depends on the position of the H-atom relative to the peroxy group, the position of other functional groups to the H-atom, and temperature (Crouse et al., 2013; Vereecken and Nozière, 2020). H-shift reactions are very slow ($k < 10^{-3} \text{ s}^{-1}$ at 298
75 K) in an aliphatic peroxy radical without a nearby electron drawing functional group for atmospheric conditions (Otkjær et al., 2018; Praske et al., 2019), so that they typically cannot compete with bimolecular reactions ($k_{\text{bi}} \sim 10^{-2} \text{ s}^{-1}$, for 50 pptv of NO and $5 \times 10^8 \text{ cm}^{-3}$ of HO_2). Even if the RO_2 radical contains a hydroxyl group, rate constants for H-shift reactions from the hydroxyl group are still low ($k \sim 10^{-2}$ to 10^{-1} s^{-1}) (Vereecken and Nozière, 2020) and can only compete with bimolecular reactions in the atmosphere in remote areas or suburban regions with low to moderate NO concentrations (50 – 500 pptv).
80 However, there are also fast H-shift reaction pathways that could compete with bimolecular reactions even in moderately or heavily polluted scenario such as allylic-H shift ($k \sim 1 \text{ s}^{-1}$) or aldehydic H-shift ($k \sim 1 - 100 \text{ s}^{-1}$) reactions, in which the radical is stabilized by delocalized electrons or the carbonyl group, respectively (Otkjær et al., 2018; Zhang and Dibble, 2011; Vereecken and Nozière, 2020). Concerning limonene, which contains two carbon-carbon double bonds (Fig. 1), recent theoretical studies suggest that the RO_2 radical formed after the reaction with OH or O_3 undergoes rapid unimolecular reactions
85 (reaction rate constants range from 1 to 10^2 s^{-1} ; Møller et al., 2020; Chen et al., 2021), which would be competitive with bimolecular reactions even in polluted environments. In this study, the atmospheric degradation of limonene was investigated in the atmospheric simulation chamber SAPHIR (Simulation of Atmospheric PHotochemistry In a large Reaction Chamber) at Forschungszentrum Jülich, Germany. Experiments were performed at three atmospherically relevant NO concentrations ranging from 0.1 to 10 ppbv. In addition, one ozonolysis experiment was conducted without the presence of NO_x . The main
90 objective of this work is to evaluate the performance of the current chemical model (Master Chemical Mechanism version 3.3.1, <http://mcm.leeds.ac.uk/MCM>) by comparing observations of trace gas and radical concentrations with model results. In particular, the question of whether the regeneration of OH from the radical cycle can be described by model calculations is investigated.



95 **Figure 1.** Simplified mechanism of the limonene oxidation by OH and O₃. If available, names of the species as they appear in the MCM v3.3.1 are given. For simplicity the stereo-specificity of the intermediates, RO₂ reactions with HO₂ and other RO₂, as well as the formation of organic nitrates are not shown here. Production of HCHO from the first oxidation step of limonene is labelled in red; HCHO production from the second oxidation step is labelled in blue; production of a HO₂ radical is labelled in orange. The production of HCHO from the decomposition of LIMALBO is not classified as it is not produced from the
 100 oxidation of the terminal C=C double bond. Most of the mechanism and branching ratios are taken from the MCM v3.3.1, except for the branching ratios of the O₃ addition at the external C=C bond of limonene (Wang and Wang, 2021) and the decomposition of LIMOOB to L5O2.



105 2 Methods

2.1 Atmospheric simulation chamber SAPHIR

SAPHIR is a large outdoor chamber located at Forschungszentrum Jülich, Germany. The cylindrical-shaped chamber (18 m length, 5 m diameter, volume 270 m³) is made of a double-wall Teflon (FEP) film, which minimizes wall loss and allows the transmission of the entire spectrum of solar radiation. A shutter system can be opened and closed to allow experiments in illuminated or dark conditions. All experiments are performed in synthetic air produced by evaporating ultrapure liquid oxygen and nitrogen (Linde, purity > 99.9999%). The air pressure inside the SAPHIR chamber is kept above the atmospheric pressure by 35 Pa to ensure that air outside the chamber cannot leak into the chamber, resulting in a typical dilution rate of trace gases of approximately 4 % h⁻¹. The temperature in the chamber is similar to the ambient temperature. More details of the chamber can be found in previous publications Rohrer et al. (2005) and Kaminski et al. (2017).

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2.2 Instrumentation

Table 1 provides an overview of the quantities that were measured and the corresponding instruments. NO and NO₂ concentrations were monitored with a chemiluminescence instrument (Eco Physics) with a blue-light photolytic converter that converts NO₂ to NO. In some experiments, the zero value of the instrument was not accurately determined as could be seen by a significant NO and NO₂ mixing ratio (NO < 0.02 ppbv, NO₂ < 0.5 ppbv) that was measured in the dark, clean chamber. For these experiments, the instrumental zero was subtracted for the analysis of data in this work. Nitrous acid was measured in some experiments by a long-path absorption photometer (Li et al., 2014). Similar as for the NO_x instruments, non-zero HONO (< 100 pptv) was measured in the clean dark chamber in some experiments. This value was attributed to an unaccounted instrumental zero value and subtracted from the data.

125 O₃ was measured with a UV absorption instrument (Ansyco). Photolysis frequencies were calculated from solar actinic flux densities measured by a spectroradiometer outside the chamber (Bohn et al., 2005; Bohn and Zilken, 2005).

OH was measured with a laser-induced fluorescence (LIF) instrument, in which OH is excited by laser light at a wavelength of 308 nm in a low-pressure cell (3.5 hPa) (Holland et al., 1995). HO₂ was measured by the LIF instrument in a separate detection cell (HO_x-cell), in which HO₂ is first converted to OH in the reaction with NO and the sum of OH and HO₂ (=HO_x) is detected by OH fluorescence. The HO₂ concentration is then calculated by subtracting the OH concentration measured in the OH-cell from the signal obtained in the HO_x-cell. Organic peroxy radicals (RO₂) can potentially interfere in the measurement of HO₂ due to the concurrent conversion of RO₂ to HO₂ after the reaction with NO. This interference is usually less than 5% for simple alkoxy radicals (C₁-C₄) because of the slow conversion rate of the alkoxy radical produced in the reaction of RO₂ with NO at the reduced O₂ concentration in the low-pressure cell. However, β-hydroxy alkoxy radicals formed from the OH-initiated oxidation of larger alkene species (e.g., isoprene) could rapidly decompose and then react with O₂ to

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form HO₂, which adds to the sampled HO₂ radical concentration (Fuchs et al., 2011; Lu et al., 2012). During the experiments in this work, the HO₂ cell was operated with low NO concentrations to minimize the conversion of RO₂ to HO₂ radicals, but it cannot be excluded that a small fraction of limonene RO₂ acted as interference.

RO₂ was measured in the RO_x-LIF system, which consists of a converter mounted on top of a fluorescence detection cell. In the converter, RO₂ is converted to HO₂ in the presence of added CO and NO at low total pressure (~ 25 hPa). The HO₂ is then passed into a low-pressure detection cell (3.5 hPa) where HO₂ is converted into OH by reaction with a large excess of added NO. The RO₂ concentration is calculated by subtracting the HO_x concentration from the measured signal (Fuchs et al., 2008, 2011). The sensitivity of the LIF instrument to RO₂ is regularly calibrated for methylperoxy radicals. Laboratory tests show that the instrument's sensitivity for RO₂ radicals from limonene might be slightly reduced compared to the sensitivity for methylperoxy radicals (0.85±0.05).

Apart from the detection by the LIF instrument, OH radical concentrations were measured by differential optical absorption spectroscopy (DOAS) (Dorn et al., 1995) in the experiments in 2015. The mean differences between the LIF-measured OH and DOAS-measured OH concentrations in the two experiments in 2015 were -13 % and +39 %. The difference between measurements on the latter experiment was higher than the combined accuracies of the measurements. Because DOAS is inherently a calibration-free method and it is regarded as a reference method (Schlosser et al., 2007, 2009), OH concentrations from this instrument are used for the evaluation of experiments in 2015.

Measurements of the OH reactivity (k_{OH}), the inverse chemical lifetime of the OH radical, were achieved by an instrument making use of laser flash-photolysis combined with the OH detection by LIF (Lou et al., 2010; Fuchs et al., 2017). In this instrument, air is sampled through a flow tube. A high concentration of OH is generated by ozone photolysis in the presence of water vapour using a short laser pulse at 266 nm, followed by the time-resolved measurement of the OH concentration while the OH is consumed by OH reactants contained in the sampled air. The pseudo-first-order rate loss constant of the decay of the OH concentration gives the OH reactivity value.

Limonene concentrations were measured by gas chromatography coupled with a flame ionization detector (GC-FID, Kaminski (2014)), as well as by a proton-transfer-reaction time-of-flight mass spectrometer (PTR-TOF-MS) (Lindinger et al., 1998; Jordan et al., 2009). Two GC-FIDs were used in the experiments in 2012. The differences between measurements by the 2 GC-derived measurements for limonene concentrations were about 12%. Discrepancies between the limonene concentrations measured by GC-FID and PTR-MS were about 12% depending on which GC-FID measurement is used.

The amount of injected limonene can also be derived from the rapid increase of OH reactivity, when limonene is injected. The injected limonene concentrations derived from the increase of OH reactivity agree with the PTRMS measurements within 15% in three experiments (08 August 2012, 10 August 2012, and 13 June 2015), but discrepancies were between 20 – 55% in the other experiments. For the analysis in this work, PTR-MS measurements were scaled in some experiments to match the



limonene injections assuming that the measurement of OH reactivity is more accurate than PTR-MS measurements because no calibration of the instrument is required.

170 A time-of-flight chemical ionization mass spectrometer (ToF-CIMS) applying ionization with nitrate ions ($^{15}\text{NO}_3^-$) detected some nitrated oxidized species in one limonene-OH oxidation experiment with high NO concentrations (12 August 2013) (Zhao et al., 2018). Although there was no calibration available to derive absolute concentrations, the time series of mass to charge ratios gave information about the relative abundance of these species.

175 Measurements of formaldehyde (HCHO) concentrations were available from either one of three instruments which included an instrument making use of the Hantzsch reaction (AL4021, Aero Laser GmbH), the DOAS instrument that is also used for the detection of OH, and a cavity ring-down spectroscopy (CRDS) analyzer (Picarro G2307) (Glowania et al., 2021). Measurements using the Hantzsch method were available in the experiments in 2012; measurements with DOAS were available in the experiments in 2015; measurements with CRDS were available in the experiments in 2019. Measurements were corrected for an unaccounted instrumental zero value (< 0.5 ppbv) observed in the clean dark chamber in some experiments.

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195 **Table 1.** Instrumentation for radical and trace-gas measurements in the chamber experiments.

Species	Method	Time resolution	1 σ precision	1 σ accuracy
OH	DOAS ^a	205 s	$0.8 \times 10^6 \text{ cm}^{-3}$	6.5 %
	LIF ^b	47 s	$0.3 \times 10^6 \text{ cm}^{-3}$	13 %
HO ₂ , RO ₂	LIF	47 s	$1.5 \times 10^7 \text{ cm}^{-3}$	16 %
OH reactivity	Laser flash photolysis + LIF	180 s	0.3 s ⁻¹	0.5 s ⁻¹
NO	Chemiluminescence	60 s	20 pptv	5 %
NO ₂	Chemiluminescence + photolytical converter	60 s	20 pptv	5 %
O ₃	Chemiluminescence	180 s	60 pptv	5 %
	PTR-TOF-MS ^c	40 s	15 pptv	14 %
Limonene	GC-FID ^d	45 min	4 – 8 %	5 %
	DOAS ^e	100 s	20 %	7 %
Formaldehyde	CRDS ^f	300 s	90 pptv	10 %
	Hantzsch ^g	60 s	25 pptv	8.6 %
Acetone	GC-FID	45 min	4 – 8 %	5 %
HONO	LOPAP ^h	300 s	3 pptv	10 %
Photolysis frequencies	Spectroradiometer	60 s	10 %	18 %

^a Differential optical absorption spectroscopy. ^b Laser-induced fluorescence. ^c Proton transfer reaction time-of-flight mass spectrometry. ^d Gas chromatography coupled with flame ionization detector. ^e Differential optical absorption spectroscopy, available for the experiments in 2015. ^f Cavity ring-down spectroscopy, available for the experiments in 2019. ^g not available in the experiments on 12 August 2013 and 05 June 2020 ^hLong-path absorption photometer.

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2.3 Limonene oxidation experiments

Before the start of an experiment, the chamber was cleaned by flushing dry, ultra-pure synthetic air through the chamber overnight to purge out trace gases that remained from previous experiments. To humidify the air at the start of the experiment, water vapor from boiling ultra-pure water (Milli-Q) was flushed into the chamber, until the relative humidity reached about 205 70%. The relative humidity gradually decreased as a result of the increasing temperature in the chamber over the course of the experiment and the dilution by the dry replenishing air. Before the injection of limonene, the chamber roof was opened to allow sunlight to irradiate the clean chamber air (zero air phase). In the illuminated chamber, small amounts of nitrous acid (HONO), formaldehyde, and acetone were formed with a rate of a few hundred pptv h⁻¹ (Rohrer et al., 2005). Therefore, the



primary source of OH and NO in most of the experiments was the photolysis of HONO leading also to a continuous increase
210 of nitrogen oxide concentrations (NO_x).

In total, seven experiments investigating limonene oxidation were performed. Chemical conditions can be divided by the NO
concentration levels (Table 2), for which the contributions of different RO_2 loss reactions varied between radical propagation
reactions (i.e., reaction with NO), radical termination channels (i.e., reactions with HO_2 and RO_2) and isomerisation reactions.

In the experiments with low NO mixing ratios of 0.1 to 0.15 ppbv (01 September 2012 (Fuchs et al., 2021a), 04 July 2019
215 (Bohn et al., 2021b)), about 3 to 4 ppbv of limonene was injected three times. Between each injection, limonene was oxidized
for about 90 to 120 minutes, so that most of the limonene reacted away before the next injection. To suppress NO
concentrations during the experiments, approximately 50 to 60 ppbv of O_3 produced by a silent discharge ozonizer (O3Onia)
was injected before opening the chamber roof.

In experiments with medium NO concentrations ranging from 0.25 to 0.4 ppbv (08 and 10 August 2012; Fuchs et al., 2021a,
220 b)), about 4 ppbv limonene was injected two hours after opening the roof. No additional trace gases were added. Measured O_3
mixing ratios increased from about 1 ppbv to 10 ppbv as a result of the photolysis of NO_2 , which was produced from the
reaction between peroxy radicals and NO.

In the experiment with the high NO concentrations (03 August 2015, (Bohn et al., 2021c)), about 15 ppbv of NO was injected
into the chamber before opening the chamber roof and 10 ppbv of limonene at later times. When most of the limonene was
225 consumed within two hours after the first injection, an additional injection of 10 ppbv of limonene was done.

Lastly, a limonene ozonolysis experiment was conducted on 05 June 2020 (Bohn et al., 2021a), in which no NO_x was present.
This experiment intended to elucidate the ozonolysis chemistry. The chamber roof was closed at all times. In the first half of
the experiment, about 4 ppbv of limonene was injected in addition to 45 ppbv O_3 . After three hours about 4.5 ppbv of limonene
was re-injected together with O_3 , so that O_3 mixing ratios reached 70 ppbv. 100 ppmv of CO was added 30 minutes before the
230 second injection to scavenge OH radicals that are produced from the limonene ozonolysis reaction.



Table 2. Experimental conditions during the limonene oxidation experiments. k_{bi} is the sum of the total loss rate of RO_2 due to bi-molecular reactions with NO , HO_2 , and other RO_2 calculated from measured concentrations. Values are averages except for limonene concentrations, for which maximum concentrations after the injection are given. In experiments with multiple limonene injections, the range of the injections is noted.

Experiment	[NO] (ppbv)	[HO ₂] (10 ⁸ cm ⁻³)	k_{bi} (s ⁻¹)	[OH] (10 ⁶ cm ⁻³)	[O ₃] (ppbv)	Limonene (ppbv)	Date
Ozonolysis	0	1.5 – 6.0	0.005 – 0.15	< 1	40 – 65	4	05 June 2020
Low NO	0.05 – 0.10	5 – 6	0.02 – 0.03	2 – 4	40 – 50	2 – 4	01 September 2012
	0.1	10	0.04	3	105	4	13 June 2015
	0.10 – 0.15	4 – 6	0.03 – 0.04	5 – 10	60	2 – 2.5	04 July 2019
Medium NO	0.2	3	0.05	2	5	4	08 August 2012
	0.3	4	0.07	4	5	4	10 August 2012
High NO	0.7	11	0.15	8	45	10	03 August 2015

2.4 Model calculations

The acquired measurements of trace gases and radicals are compared against a zero-dimensional box model applying the Master Chemical Mechanism (MCM) version 3.3.1 (Saunders et al., 2003; Jenkin et al., 2015). In addition to the chemistry from the MCM, chamber-specific processes including dilution and small productions of HONO, acetone and formaldehyde in the presence of sunlight are included in the model. Dilution in the chamber is implemented as a first-order loss process. The rate constant is calculated based on the monitored replenishment flow rate. The parameterization for the production of chamber sources for nitrous acid, formaldehyde and acetone follows the description in Rohrer et al. (2005); and Kaminski et al. (2017), in which production rates are parameterized as functions of temperature, relative humidity, and radiation. The source strength for the production rates are scaled for each experiment from the observed increase in concentrations for the part of the experiment when the chamber roof was opened, but limonene was not present (zero-air phase).

In addition, a background OH reactivity in the range of 1 s⁻¹ is present in the illuminated, clean chamber due to the presence of unmeasured OH reactants. In order to account for this background reactivity, an artificial OH reactant that behaves like CO is implemented in the model. Its concentration is adjusted in the model to match the observed OH reactivity during the zero-air phase and is assumed to be constant during the rest of the experiment.

The OH yield from limonene ozonolysis in the model calculations in this work is updated based on the IUPAC recommendations (Cox et al., 2020) decreasing the value from 87% to 66%. This is supported by multiple experimental studies (Aschmann et al., 2002; Herrmann et al., 2010; Forester and Wells, 2011) as well as theoretical studies (Wang and Wang,



260 2021). The decomposition product of the Criegee intermediate (MCM name: LIMOOB) is also updated, in which the terminal
carbonyl group remains during the vinyl hydroperoxide mechanism (VHP). In addition, instead of forming a primary RO₂
(MCM name: C923O2) by eliminating a carbonyl group, decomposition of LIMOOB leads to the production of a secondary
β-oxo substituted RO₂ (L5O2, Fig. 1). This is supported by the theoretical investigation of the ozonolysis of cyclohexene that
the departure of the carbonyl group is not competitive without any inducing functional group (e.g., β-hydroperoxyl) (Rissanen
265 et al., 2014).

Three simulation model runs are performed for each experiment. In all model runs, physical parameters including temperature,
pressure, photolysis frequencies, and the dilution rate of trace gases due to the replenishment flow are constrained to
observations. If available, HONO concentrations are prescribed as measured to constrain the production from the chamber. In
addition, ozone concentrations in the model are constrained to ensure that shortcomings of the model to predict ozone do not
270 complicate the interpretation by an inappropriate fraction of limonene reacting with ozone. Time series of NO and NO₂
concentrations are constrained to observations in the model runs except for the third set of simulations to avoid potential
impacts of shortcomings of the model to describe these species.

In the first set of simulations (denoted as “reference run”), model runs are performed without any further constraints or
modifications. In the second set of simulations (denoted as “constrained run”), HO₂ concentrations are prescribed as measured
275 to constrain the OH-production rate from the HO₂+NO reaction. In this model run, OH reactivity is also adjusted to the
measurements to yield an OH loss rate as observed (Section 3.3.1 to 3.3.4). In the third set of simulations (Section 3.5), NO
and NO₂ concentrations are not constrained to investigate the fate of nitrogen oxides, radical concentrations and OH reactivity
are also free parameters in the model.

OH reactivity from oxidation products tends to be overestimated by the model. Potential reasons include wall loss reactions of
280 low volatility compounds such as organic nitrates and peroxides, uncertainties from the reactivity of products and intermediates
that are mostly derived using structure-activity relationship (SAR) in the model (Jenkin et al., 1997; Saunders et al., 2003),
and chemical loss due to reactions that are not included in the MCM (e.g., isomerisation reactions). To reduce the OH reactivity
in the model, it is assumed that organic nitrates and peroxides derived from limonene oxidation undergo a first order loss
process.

285 The additional loss rate for organic nitrates is estimated from the decaying signal at the corresponding mass/charge ratio of the
least oxidized C10 nitrate species (C₁₀H₁₇NO₆) observed by the chemical ionization mass spectrometry (CIMS) instrument
after the chamber roof was closed in the experiment with high NO concentrations. The lifetime of the least oxidized species
C10 nitrate is used to estimate the loss rate, as the only C10 nitrate species from limonene oxidation that the MCM includes is
C₁₀H₁₇NO₄. The loss rate constant is equivalent to a lifetime of about 2 hours (Supplementary materials Fig. S1, Zhao et al.
290 (2018)), which is comparable to the chemical loss rate in the reaction with OH for conditions of these experiments (chemical
lifetime of 0.5 to 4 hours). The specific reason for the additional loss of nitrates could not be identified, but might be due to



wall loss or due to loss by hydrolysis reactions in the humidified air similar to findings in an isoprene-rich forest (Romer et al., 2016).

Organic peroxides, however, could not be detected by the CIMS instrument and hence the additional loss rate is adjusted such that the simulated OH reactivity agrees with measurements in the experiments with low and medium NO, when there should have been significant production of organic peroxides. This requires a lifetime of 10 minutes, much shorter than the lifetime of the order of hours under typical atmospheric conditions or those in the current experiments.

3. Results

3.1 Product yields from the oxidation of limonene

3.1.1 Formaldehyde and acetone yields from the oxidation of limonene by OH and O₃

Formaldehyde (HCHO) is one of the oxidation products of the reaction of limonene with the OH radical as well as with ozone. It is mainly produced from the oxidation of the terminal double bond in limonene (Fig. 1). The product yield of HCHO from the oxidation of limonene by OH and O₃ can be determined from the chamber experiments by comparing the measured concentrations of HCHO to the amount of oxidized limonene calculated from the measured limonene, OH and O₃ concentrations. As HCHO has also a small chamber source (Section 2.3) and it is lost by its reaction with OH and by photolysis, its concentration needs to be corrected for these losses following the method presented in Kaminski et al. (2017) and Rolletter et al. (2019).

For the evaluation of the HCHO yields, only measurements from the first two hours of the experiment after the first limonene injection are used, because further oxidation of accumulating organic products could contribute to the production of HCHO at later times of the experiment. This is demonstrated by a rapid increase of HCHO when 80% of limonene had reacted away after a roughly linear relationship between HCHO and limonene concentrations (Fig. 1 and Supplementary Material Fig. S2). To derive the HCHO yield, a linear regression is performed for data until 40% of limonene is consumed and the slope of the linear regression gives the HCHO yield. This results in values of (12±3)%, (13±3)%, and (32±5)% for the OH oxidation experiments with low, medium and high NO mixing ratios, respectively (Fig. 2).

Formaldehyde is only produced from one of the three major RO₂ species formed in the initial reaction of limonene with OH. The structure-activity relationship (SAR) (Peeters et al. 2007) gives a yield of the formation of this RO₂ radical (Fig. 1) that results from the OH addition to the terminal C=C double bond (MCM name: LIMCO₂) of 37% with an error of 15%. When the LIMCO₂ radical reacts with NO or RO₂ radicals, an alkoxy radical is formed that subsequently forms HCHO from its decomposition together with a carbonyl compound (MCM name: LIMKET, Fig. 1). Therefore, the expected HCHO yield



would be 37% at maximum from which the fraction of RO₂ reacting with other reaction partners such as HO₂ and the formation of organic nitrates in the reaction with NO (value determined in this work: (34±5) %, Section 3.1.2) needed to be subtracted. The uncertainty in the yield of organic nitrates could be higher because the values are determined for all RO₂ species from limonene oxidation by OH, but the nitrate yield for individual RO₂ could vary. The fraction of RO₂ reacting with NO is more
325 than 85% in the experiments with high and medium NO, and 65% in the experiment with low NO. The relative error of the fraction of RO₂ reacting with NO could be up to 35% due to the uncertainties of the reaction rate constants. Therefore, the HCHO yields expected from SAR in the experiments with high and medium NO is (22±9) % and in the experiment with low NO is (16±6) %.

All observed HCHO yields agree with SAR-derived HCHO yields which carry high relative uncertainties (40%). This could
330 explain the discrepancies between the observed yield for the experiment with high NO mixing ratios and the SAR-derived yield. The largest discrepancy could be observed for medium NO levels. A higher yield of LIMCO₂ or lower yield of organic nitrates for the reaction of LIMCO₂ with NO could explain the discrepancies observed for the experiments with low NO. Another possible reason is that there are other reaction channels of the LIMCO₂ peroxy radical that are competitive with the bimolecular reactions with NO and HO₂ (k_{bi} up to 0.07 s⁻¹) for conditions of the experiments with low and medium NO, but
335 would not be relevant in the experiment with high NO mixing ratios (k_{bi} up to 2 s⁻¹). These reactions channels would not produce formaldehyde such as isomerisation. Possible isomerisation reactions of LIMCO₂ are further discussed in Section 4.3.1.

The formation of HCHO at a later time in the experiments includes the production from further oxidation of first generation products. Therefore, HCHO yields are increasing over the course of the experiment to values of 40 to 90% with lowest yields
340 in the experiments with low NO (Fig. S2). Part of the differences between the experiments could be due to the differences in the contributions of various product species, so that numbers are not necessarily comparable. The HCHO yield remains below 100% in all experiments. This is consistent with HCHO being only produced from the oxidation of the terminal C=C bond of the limonene structure on a short time scale of several hours in typical atmospheric conditions, so that total HCHO yield is limited to one in the limonene oxidation scheme.

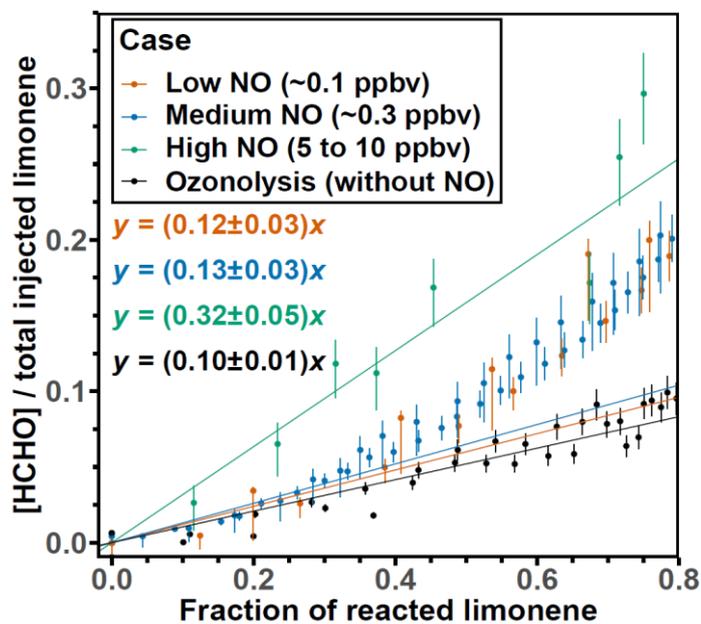
345 Previous studies reported formaldehyde yields for the OH-oxidation of limonene between 36 and 43% (Larsen et al., 2001; Lee et al., 2006) (Table 3). The experiments by Larsen et al. (2001) and Lee et al. (2006) were performed at high NO concentrations, but the yield of HCHO yield is higher than the HCHO yield determined in the experiment with high NO in this work. However, experiments by Larsen et al. (2001) and Lee et al. (2006) were performed with much higher concentrations of limonene and NO (Table 3) and it is not clear, if only HCHO from the first oxidation step is considered (Fig. 1), so that
350 numbers may not be entirely comparable. In the study by Librando and Tringali (2005), formaldehyde yields were determined to be 27 and 92% in experiments with high limonene mixing ratios of 13 and 2 ppmv, respectively, but in the absence of NO. Yields in their experiments may not be comparable to yields in this work, because the fate of organic peroxy radicals was likely dominated by atmospherically not relevant RO₂+RO₂ recombination reactions.



The HCHO yield derived from the pure ozonolysis experiment in this work is $(10\pm 1)\%$ in the presence of the OH scavenger.
355 The same value is obtained, if the formaldehyde yield is determined from the part of the ozonolysis experiment without OH
scavenger, when approximately 40% of limonene reacted with OH that is produced from the ozonolysis reaction. Similar to
the OH reaction, formaldehyde production is expected from the subsequent chemistry after the ozone addition to the terminal
C=C double bond. Because ozone preferably (87%) adds to the endocyclic C=C double bond (Wang and Wang, 2021), a low
yield can be expected. As there was no NO present during the experiment, production of formaldehyde from the reaction
360 between limonene and OH could only be possible from the reaction of LIMCO₂ with peroxy radicals or unimolecular reactions
(e.g. isomerisation). As the HCHO yield is the same independent on the addition of the OH scavenger, the yield from the
limonene-OH reaction at zero NO must be similar to the HCHO yield from the limonene ozonolysis.

The formaldehyde yield from the limonene ozonolysis derived in this work agrees well with the yield of 10% determined in
the work by Grosjean et al. (1993) performed with an OH scavenger. Yields were also determined in the work by Gong et al.
365 (2018) from several experiments at various chemical conditions. Values of 5 to 11% derived in their experiments with a
limonene:O₃ concentration ratio of 1:2 are consistent with results in this work. However, the HCHO yield was 11 to 27% in
the experiments with a limonene:O₃ concentration ratio of about 1:100. This higher yield might be explained by additional
formaldehyde production from the ozonolysis of secondary products (Fig. 1). It is worth noting that no significant differences
could be observed in experiments with and without OH scavenger in the experiments by Gong et al. (2018) similar to findings
370 in this work.

In the experiments in this work, there was no significant acetone production from the oxidation of limonene, as the production
rate of acetone from the chamber source (~ 80 pptv h⁻¹) could already explain the observed increase of the acetone concentration
during the experiments. This suggests that unlike other monoterpenes, acetone is not a significant product from limonene
oxidation. This is consistent with findings by Lee et al. (2006) and Larsen et al. (2001). Also, the MCM does not predict
375 acetone as a product in the limonene oxidation scheme.



380 **Figure 2.** HCHO concentrations divided by the injected limonene concentration plotted versus the fraction of reacted limonene for the first injection in the experiments with different NO levels. The regression includes data points up to 40% of the fraction of reacted limonene.

385

390



Table 3. HCHO yields from the OH- and O₃-oxidation of limonene determined in the experiments in this work and reported in literature. Maximum values for limonene, O₃, and NO concentrations in the experiments are listed.

Reaction	HCHO yield (%)	Limonene (ppbv)	NO (ppbv)	O ₃ (ppbv)	OH scavenger
Limonene + OH					
Larsen et al., 2001	36 ± 5	1000	1000		
Librando and Tringali, 2005	27 – 92	2100 – 13200	0		
Lee et al., 2006	43 ± 5	120	132 ^a		
This work	12 ± 3	3.5	0.05 – 0.15		
	13 ± 3	3.5	0.3		
	32 ± 5	10	3 – 10		
Limonene + O₃					
Grosjean et al., 1993	10	1200	70 – 100		Yes ^b
Gong et al., 2018	7 – 11	280	500		No
	5 – 8	280	500		Yes ^c
	13 – 27	183	19000		No
	11 – 23	183	19000		Yes ^c
This work	10 ± 1	4	50		No
	11 ± 1	4	80		Yes ^d

^aGiven as NO_x concentration ^b200 ppmv cyclohexane ^c400 ppmv of 2-butanol or cyclohexane ^d60 ppmv CO

395 3.1.2 Organic nitrate yield from the oxidation of limonene by OH

The organic nitrate yield of the RO₂+NO reaction for peroxy radicals produced from the limonene+OH reaction is determined for experiments in this work following the method described for the analysis of previous experiments in the SAPHIR chamber (Hantschke et al., 2021; Tan et al., 2021). The amount of organic nitrates (RONO₂) in the chamber is estimated by the difference between the accumulated production of nitrogen oxides from the chamber source of HONO (Q_{HONO}), which is the only relevant
 400 source for nitrogen oxides in the experiments, and the total NO_y (=NO_x + HONO + HNO₃) concentrations. This method assumes that the differences can be solely attributed to the production of peroxy acyl nitrate (PAN)-like species, which can act as a reservoir for nitrogen oxides and can be significant in the oxidation of specific organic compounds.

The source strength of HONO can be calculated from the photo-stationary state between HONO, NO and OH using measured concentrations and measured photolysis frequencies:

$$405 \quad \frac{d[\text{HONO}]}{dt} = Q_{\text{HONO}} - j_{\text{HONO}}[\text{HONO}] + k_{\text{OH}+\text{NO}}[\text{OH}][\text{NO}] \approx 0 \quad (1)$$



$$Q_{\text{HONO}} = j_{\text{HONO}}[\text{HONO}] - k_{\text{OH}+\text{NO}}[\text{OH}][\text{NO}] \quad (2)$$

The total amount of organic nitrates can be calculated from the difference between the accumulated HONO production and the sum of other nitrogen oxide species considering also their loss due to dilution in the chamber with rate k_{dil} and the loss of NO_2 in the reaction with OH (Hantschke et al., 2021):

$$410 \quad [\text{RONO}_2] = \int (Q_{\text{HONO}} - k_{\text{OH}+\text{NO}_2}[\text{NO}_2][\text{OH}] - k_{\text{dil}}([\text{NO}] + [\text{NO}_2] + [\text{HONO}])) dt - ([\text{NO}] + [\text{NO}_2] + [\text{HONO}]) \quad (3)$$

Assuming that at least right after the injection of limonene, the production of organic nitrates is dominated by the reaction of NO with RO_2 derived from limonene-OH, the organic nitrate yield $\Phi_{\text{limonene}+\text{OH}}$ can be calculated:

$$[\text{RONO}_2] = \Phi_{\text{limonene}+\text{OH}} \int ([\text{RO}_2]_{\text{limonene}} \times [\text{NO}] \times k_{\text{RO}_2+\text{NO}}) dt \quad (4)$$

415 The fraction of RO_2 from limonene-OH oxidation to the measured total RO_2 concentration is estimated based on the ratio of the OH reactivity from limonene to the OH reactivity from all OH reactants that are expected to produce RO_2 in the reaction with OH (Tan et al., 2021). The latter can be calculated by subtracting the background OH reactivity ($k_{\text{OH}_{\text{bg}}}$) that includes the reactivities attributed to inorganic species and formaldehyde from the total measured OH reactivity (k_{OH}):

$$[\text{RO}_{2_{\text{limonene}}}] = \frac{k_{\text{OH}_{\text{limonene}}}}{k_{\text{OH}} - k_{\text{OH}_{\text{bg}}}} [\text{RO}_2] \quad (5)$$

Using Equation 3 and 4, the nitrate yield $\Phi_{\text{limonene}+\text{OH}}$ is obtained as the slope of the linear regression between the calculated
420 organic nitrate concentrations and the integrated turnover rate of the reaction between $\text{RO}_{2_{\text{limonene}}}$ and NO. Only experiments with medium NO mixing ratios are used for this analysis here, as HONO measurements were performed and more than 90% of the limonene was oxidized by OH in these experiments. This results in a $\Phi_{\text{limonene}+\text{OH}}$ of $(34 \pm 5)\%$ (Fig. 3). The uncertainty of this calculation is mainly due to the accuracy of concentration measurements of NO_y species and radicals. The organic nitrate yield of the RO_2 from limonene ozonolysis is not determined in this study as a considerable amount of PANs could be
425 formed according to the MCM model. Therefore the organic nitrate yield analysis is not conducted for the experiments at low NO mixing ratio.

The organic nitrate yield in this study is higher than values estimated using different SAR approaches of 19% (Jenkin et al., 2019) to 28% (Arey et al., 2001; Leungsakul et al., 2005). However, the value agrees well with nitrate yield derived in experiments, in which the yield was derived from the analysis of the aerosol chemical composition produced from the oxidation
430 of limonene ($(36 \pm 6)\%$; Rollins et al. (2010)).

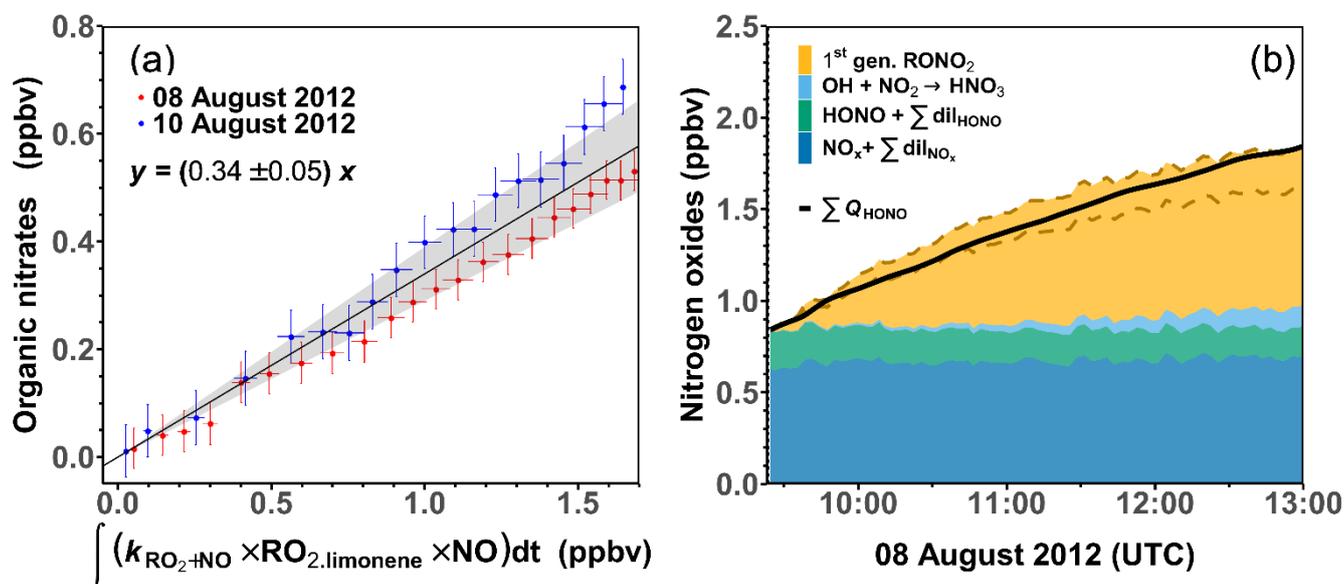


Figure 3. (a) Regression between calculated organic nitrate mixing ratio and integrated turnover rate of the reaction of limonene-derived RO₂ and NO. The regression only includes data when less than 60% of the limonene from the first injection reacted away. (b) Concentrations of nitrogen oxide species after the first limonene injection in the experiment with medium NO on 08 August 2012 compared to the total production of HONO from the chamber source. The organic nitrate contribution is calculated using a nitrate yield of (34±5)% for the reaction of limonene RO₂ with NO, the uncertainty range is enclosed by the two dashed lines.

3.2 Comparison of modelled and measured concentration time series

3.2.1 Model-measurement comparison for limonene photo-oxidation experiments at low NO mixing ratio

Three experiments were performed at low NO concentrations on 01 September 2012 (Fig. 4), 13 August 2015 and 04 July 2019 (Supplementary Material, Fig. S3 and S4). Measured OH concentrations during the limonene oxidation were about (2-7)×10⁶ cm⁻³ in the experiments in 2012 and about (5-11)×10⁶ cm⁻³ in the experiments in 2019. OH concentrations were highest, when most of the limonene was consumed. The overall higher OH concentrations in the experiments in 2019 were due to the higher NO mixing ratios (~0.15 ppbv) compared to the experiment in 2012 (~0.05 ppbv), which led to a faster OH production from the radical regeneration reaction of HO₂ and NO. In these experiments, around 30 to 40% of limonene reacted with O₃ and 60 to 70% of limonene reacted with OH. Measured HO₂ concentrations were around (5-8)×10⁸ cm⁻³. Measured RO₂ concentrations were around (2-5)×10⁸ cm⁻³ in the experiment in 2012 and (5-10) ×10⁸ cm⁻³ in the experiments in 2019, respectively. The average chemical lifetime of RO₂ to bimolecular reactions with NO, HO₂ and RO₂ can be calculated using



measured NO, HO₂ and RO₂ concentrations, respectively, resulting in a chemical lifetime of 25 to 50s. 60 to 70% of the RO₂ radicals reacted with NO and 30 to 40% reacted with HO₂. Loss due to RO₂+RO₂ recombination reactions are predicted not to play a major role, if reaction rate constants are taken from the MCM that are much lower compared to the rate constants of the reactions with HO₂.

455 In the reference model run, which essentially uses the MCM chemistry without constraining radical production or destruction processes (Section 2.4), the OH reactivity is overestimated in the model by 3 to 5 s⁻¹ after nearly all limonene has reacted away. Right after the limonene injections, HO₂ concentrations are underestimated by 20 to 50% and the RO₂ concentrations are overestimated by 500 to 700%. OH is underestimated by at least 50%, which could partly result from the overestimation of the OH reactivity and underestimated HO₂ concentration (and thereby the reaction rate of HO₂+NO). The underestimated OH
460 concentration also leads to a slightly slower decay of the modelled limonene, so that modelled limonene concentrations are about 0.1 – 0.2 ppbv higher than measured values at the end of the experiment.

In the model run with adjusted OH reactivity and constrained HO₂ concentration (constrained model run), the discrepancy between modelled and measured OH concentrations is reduced. They are still underestimated by 33% on average throughout the whole experiment and by about 50% one hour after the limonene injection, which could indicate a missing OH production
465 process in the model. However, the temporal behaviour of the modelled limonene concentration agrees with the behaviour of the measurements, suggesting that measured OH concentrations are too high.

Maximum modelled RO₂ concentrations are reduced by about 30% compared to the reference run as a result of the increased HO₂ concentration in the constrained run leading to a higher RO₂ loss by RO₂+HO₂ recombination reactions. However, this improvement is insufficient to explain measured RO₂ concentrations, which are still at least a factor 4 lower than modelled
470 values.

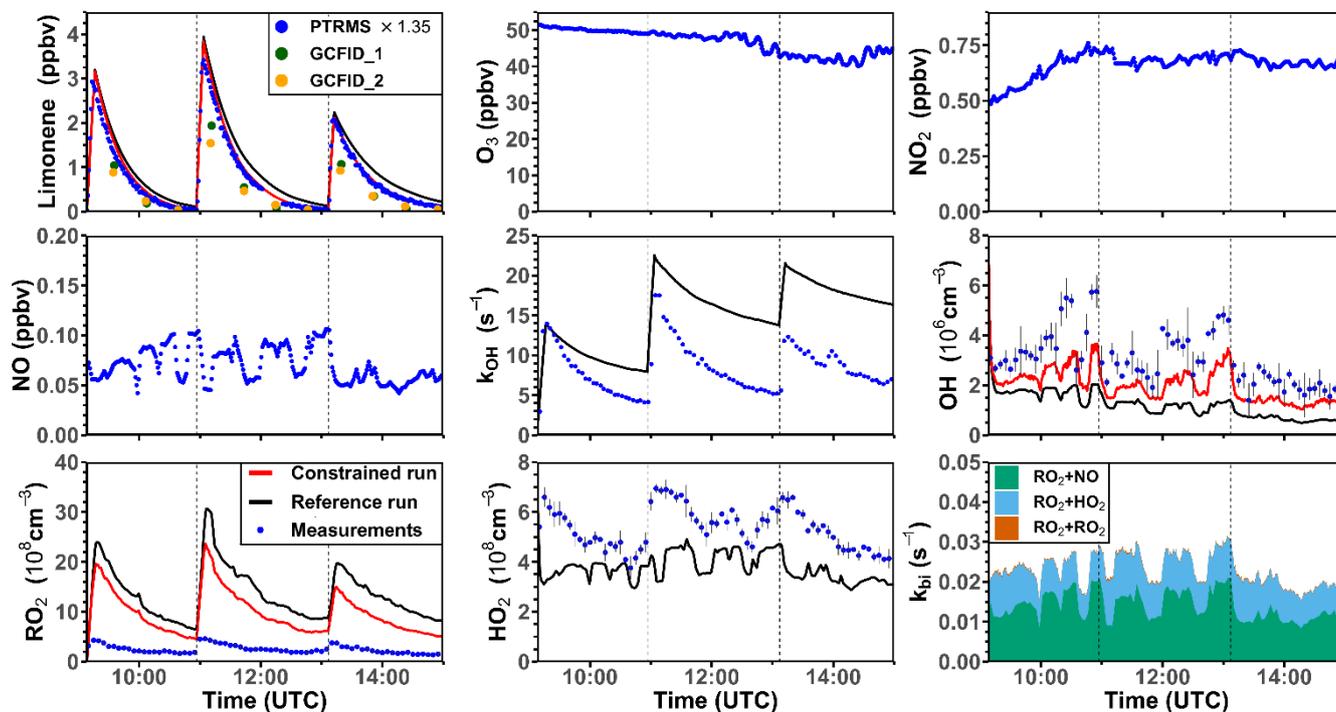


Figure 4 Time series of radicals, inorganic and organic species in the limonene oxidation experiment with low NO mixing ratios on 01 September 2012. Limonene concentrations measured by the PTRMS are scaled by a factor of 1.35 to match the increase of OH reactivity during the injections of limonene. The vertical dashed lines represent times, when limonene was injected into the chamber. RO₂ bi-molecular reaction loss rates (k_{bi}) are calculated based on the measured NO, HO₂, and RO₂ concentrations using the reaction rate constants as included in the MCM. In the constrained model run HO₂ concentrations are constrained to measurements and the OH reactivity is adjusted by additional loss reactions to match measurements.

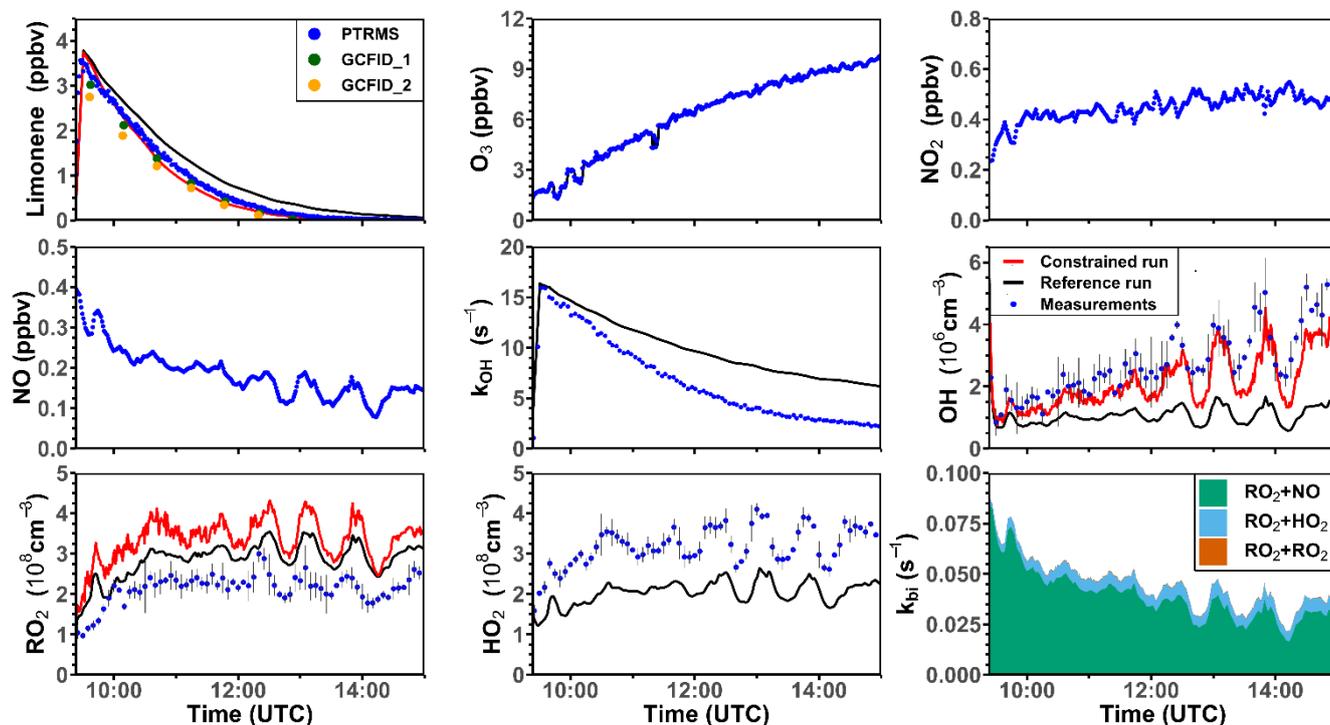
3.2.2 Model-measurement comparison for limonene photo-oxidation experiments at medium NO mixing ratios

Two oxidation experiments were performed at medium NO levels on 08 and 10 August 2012 (Fig. 5 and Supplementary Material Fig. S5). Throughout both experiments, measured NO mixing ratios ranged from 0.25 to 0.4 ppbv. More than 90% of the RO₂ derived from the limonene-OH reaction reacted with NO. O₃ mixing ratios increased gradually from about 1 ppbv to 10 ppbv as a result of the photolysis of NO₂, which was produced from the reaction between organic peroxy radicals and HO₂ with NO. At such low O₃ concentrations, the reaction with OH was the dominant loss pathway for limonene (90%). The measured OH concentration ranged from (2-5)×10⁶ cm⁻³ and the measured HO₂ and RO₂ concentrations were about (2-5)×10⁸ cm⁻³.



In the reference model run, the simulated OH reactivity increasingly deviates from measurements over the course of the experiment. The differences are about 2 to 4 s⁻¹ after two to three hours of oxidation when all limonene has been consumed. Contributors to the OH reactivity in the model in addition to limonene are mostly oxidation products such as aldehydes and ketones (~ 70%), organic peroxides (~ 15%) and organic nitrates (~ 15%) (Supplementary Material, Fig. S6). Simulated HO₂ concentrations are underestimated in the reference model run by around 10 to 30%. Hence, the simulated OH concentration is underestimated by 30 to 80% due to the slow regeneration rate from the reaction between HO₂ and NO as well as the faster removal rate of OH. As a result, limonene concentrations are overestimated by the model up to 0.3 ppbv throughout the experiment.

In the constrained model run, the simulated OH concentration is on average 20% lower than measured values. This difference is within the uncertainties of OH measurements. The higher modelled OH concentrations compared to the reference model run leads to a faster consumption of limonene compared to the reference model. This temporal behaviour better agrees with the temporal behaviour observed by PTRMS measurement, but the limonene consumption is slightly faster in the model than measurements suggests. Modelled RO₂ concentrations are 50 to 100% higher than measurements. Values are 10 to 20% higher compared to the reference model run right after the injection of limonene presumably due to the enhanced RO₂ production in the constrained model run, in which OH concentrations are higher.



505

Figure 5. Time series of radicals, inorganic and organic species in the limonene oxidation experiment at medium NO level on 08 August 2012. RO₂ bi-molecular reaction loss rates (k_{bi}) are calculated based on the measured NO, HO₂, and RO₂ concentrations using the reaction rate constants from the MCM. In the constrained model run HO₂ concentrations are constrained to measurements and the OH reactivity is adjusted by additional loss reactions to match measurements.

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3.2.3 Model-measurement comparison for limonene photo-oxidation experiment at high NO mixing ratios

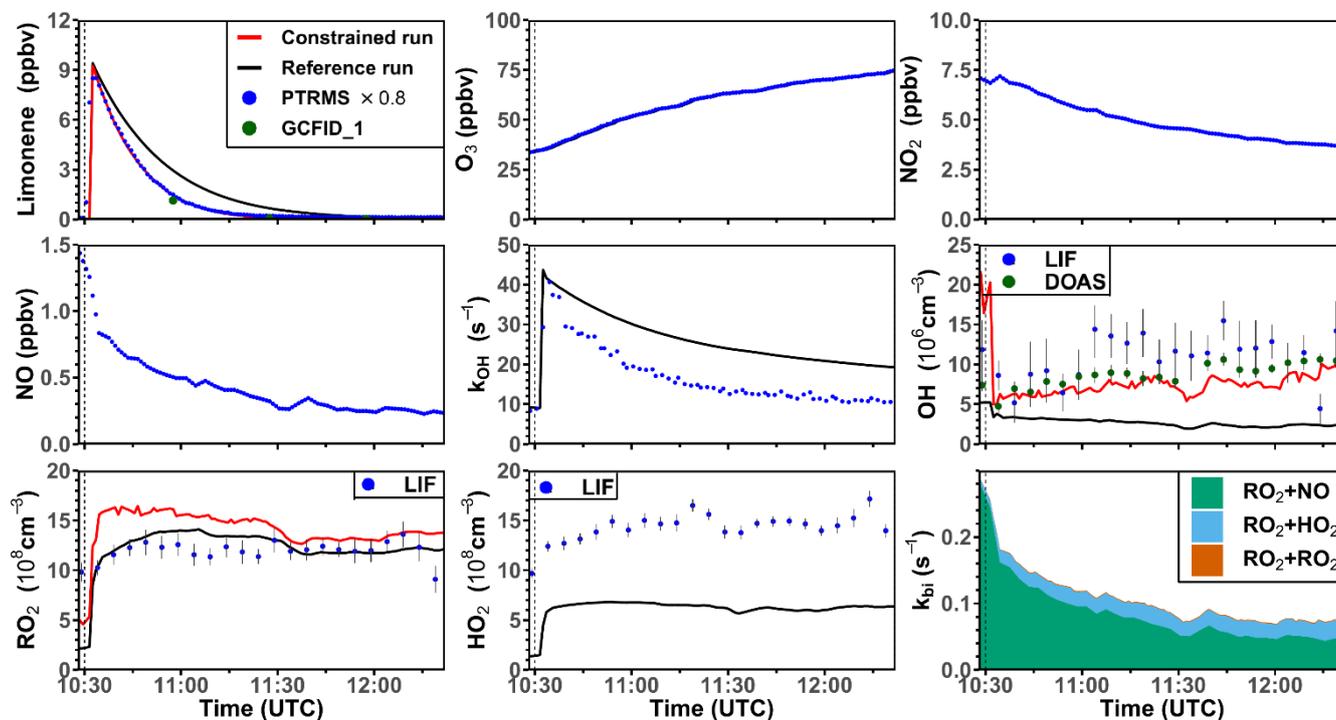
One experiment was performed with high NO mixing ratios on 03 August 2015. In this experiment, limonene was injected twice. Unfortunately, measurements of radicals failed during the first part of the experiment, so that only the part of the experiment after the second limonene injection can be analysed (Fig. 6). The ozone mixing ratio at the beginning of this part of the experiment was about 35 ppbv as a result of ozone production during the first part of the experiment. The OH concentration measured by DOAS was around $(5-10) \times 10^6 \text{ cm}^{-3}$, so that around 70% and 30% of limonene reacted with OH and O₃, respectively. The OH concentration measured by LIF was around $(5-15) \times 10^6 \text{ cm}^{-3}$. The HO₂ concentration was about $1 \times 10^9 \text{ cm}^{-3}$ throughout the experiment. The high NO and HO₂ concentrations led to a short RO₂ lifetime of 5 to 10 s, much shorter than in the other experiments with lower NO concentrations. The removal of RO₂ by bimolecular reactions was mostly due to the reaction with NO (85 %) and HO₂ (15 %).

520



In the reference model run, the model-measurement discrepancy of the OH reactivity is around 3 s^{-1} before the second limonene injection, and it further increases to 10 s^{-1} at the end of the experiment. The absolute discrepancy in OH reactivity is higher compared to other experiments because the total amount of limonene and therefore the production of oxygenated products was higher. Modelled OH and HO₂ concentrations are both lower by 50 to 70% than the measured values. The low OH
525 concentration in the model again leads to a slower decay of modelled limonene concentrations compared to observed values. RO₂ concentrations show a good agreement with the measurement, with a discrepancy of less than 20%.

In the constrained model run, the simulated OH concentration is in good agreement with the measurement by DOAS in the first hour and only slightly underestimates the measurements by about 20% at the end of the experiment. As a result, the agreement between modelled and measured time series of limonene concentrations improves compared to the reference model
530 run. Differences between RO₂ concentrations in the constrained and reference model are rather small except for the point in time when limonene is injected and RO₂ concentrations in the constrained model run increase more rapidly than observations. This results in 30% difference between modelled and measured values during the first hour of the experiment. This can be explained by the increased RO₂ production from the limonene oxidation by the OH radical which concentration is underestimated in the reference model run. Possible reasons for the overestimated RO₂ concentration could include an
535 underestimated loss rate of RO₂ at NO mixing ratio of about 1 ppbv (Section 3.4), or an overestimated production of RO₂ from the further oxidation of products from the previous limonene injection.



540 **Figure 6.** Time series of radicals, inorganic and organic species during the limonene oxidation experiment at high NO mixing ratios on 03 August 2015. Limonene concentrations measured by the PTRMS are scaled by a factor of 0.8 to match the increase of OH reactivity during the injection of limonene. RO₂ bi-molecular reaction loss rates (k_{bi}) are calculated based on the measured NO, HO₂, and RO₂ concentrations using the reaction rate constants from the MCM. In the constrained model run HO₂ concentrations are constrained to measurements and the OH reactivity is adjusted by additional loss reactions to match
545 measurements.

3.2.4 Model-measurement comparison for limonene ozonolysis experiment in the dark with and without OH scavenger

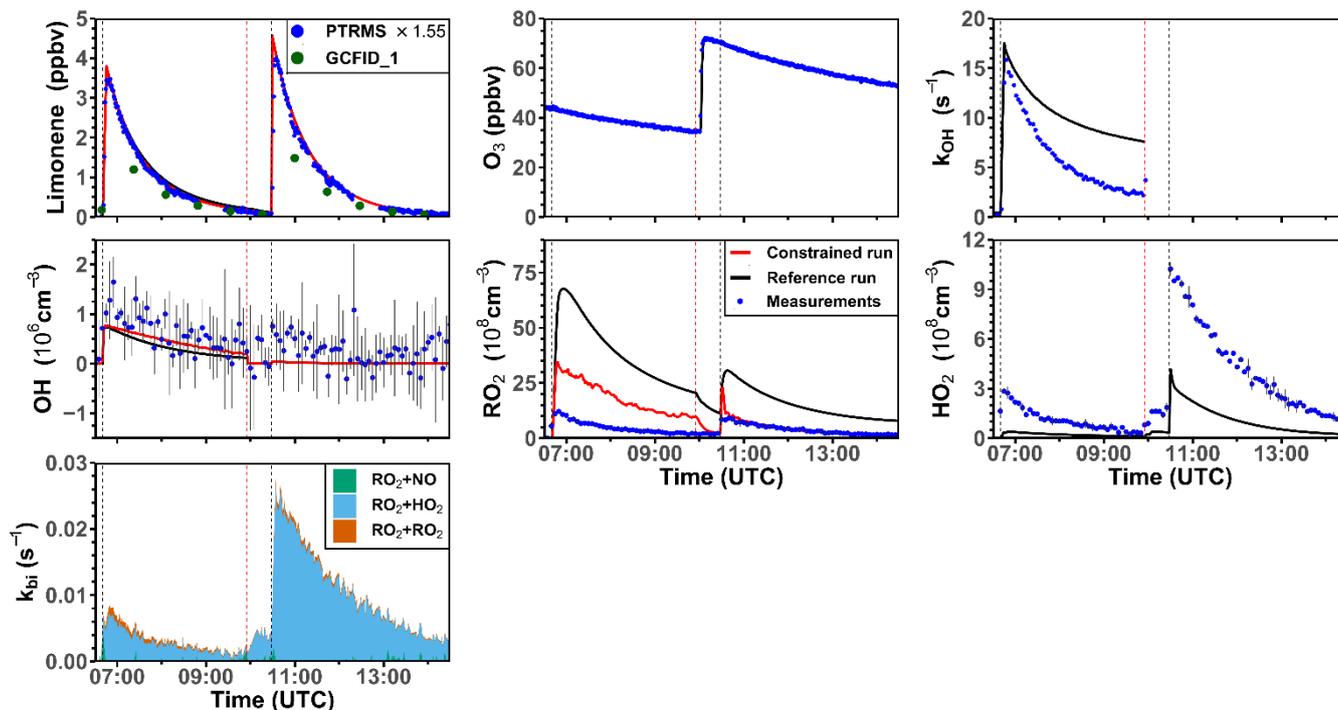
The ozonolysis of limonene in the dark was investigated in a separate experiment on 05 June 2020 (Fig. 7). A significant OH concentration was observed with values up to $1.2 \times 10^6 \text{ cm}^{-3}$ after the first injection of limonene as a result of the OH production
550 from the ozonolysis reaction. OH concentrations were below the limit of detection of both DOAS and LIF after 100 ppmv CO had been injected as OH scavenger right before the second limonene injection. The measured HO₂ concentration was around $3 \times 10^8 \text{ cm}^{-3}$ after the first limonene injection and increased to maximum concentrations of $1 \times 10^9 \text{ cm}^{-3}$ after the second limonene injection due to the conversion of OH radicals formed from the ozonolysis in the reaction with excess CO. Because there was no limonene oxidation by OH, there was also less production of RO₂ radicals than the first limonene injection. However, the
555 observed maximum RO₂ concentration after each limonene injection was about $1 \times 10^9 \text{ cm}^{-3}$.



Similar as observed in the other experiments, modelled OH reactivity is higher than measured values by 5 s^{-1} . The maximum discrepancies appear right before the injection of CO, after which OH reactivity could not be measured due to the high contribution of CO. In the reference model run, HO_2 is underestimated by the model by around 90% during the first part of the experiment and modelled HO_2 is also a factor of three lower than measurements during the pure ozonolysis part in the presence of CO as OH scavenger. In contrast, measured RO_2 concentrations are drastically overestimated by the model by a factor of 7 right after each limonene injection.

In the constrained model run, simulated OH concentrations are 25% higher compared to the reference model as a result of the reduced OH reactivity in the model. The main difference between the results of the reference model run and the constrained model run is that RO_2 concentrations are reduced by half in the constrained model. This is caused by an increased RO_2 removal rate by the high HO_2 concentration.

Using measured RO_2 and HO_2 concentrations, the maximum reaction rates of the reaction of RO_2 with HO_2 is about 0.005 s^{-1} and 0.02 s^{-1} for the first and second limonene injection, respectively. In contrast, the RO_2 loss rate due to the recombination of organic peroxy radicals is calculated to be only 0.001 s^{-1} , so that it is expected that RO_2 reacted mainly with HO_2 to form peroxides, if only bimolecular reactions are considered. However, the loss due to RO_2 - RO_2 reactions could be of similar importance as the loss due to RO_2 - HO_2 reactions, if low HO_2 and high RO_2 concentrations as predicted by the reference model run are used for the calculation. The large discrepancies between measured and modelled RO_2 and HO_2 radical concentrations and possible explanations are further discussed in Sections 4.1 and 4.2.



575 **Figure 7.** Time series of radicals, inorganic and organic species during the limonene ozonolysis experiment in the dark on 05
June 2020. Limonene concentrations measured by the PTRMS are scaled by a factor of 1.55 to match the increase of OH
reactivity during the injection of limonene. The black vertical dashed lines represent when limonene was injected into the
chamber; the red vertical dashed line represents the injection of 100 ppmv of CO. RO₂ bi-molecular reaction loss rates (k_{bi}) are
580 calculated based on the measured NO, HO₂, and RO₂ concentrations using the reaction rate constants in the MCM. In the
constrained model run HO₂ concentrations are constrained to measurements and the OH reactivity is adjusted by additional
loss reactions to match measurements.

3.3 Chemical OH radical budget using measured quantities

To evaluate whether unaccounted chemical processes are required to explain the total OH destruction rate in the experiments,
585 the rates of OH-producing reactions are calculated using measured trace gas and radical concentrations. Turnover rates of
single reactions are summed up and compared to the total OH destruction rate. Due to the short chemical lifetime of OH
radicals (< 1 s) steady-state conditions are expected that the production and destruction rates must be balanced.

The total OH destruction rate is obtained by computing the product of the measured OH reactivity and OH concentration. In
experiments in year 2015, when OH was measured by both DOAS and LIF instruments, DOAS measurements are used for the



590 calculation of total OH destruction rate. The uncertainty of the OH destruction rate is 16% calculated by error propagation of
the measurement uncertainties. The main OH production processes included in the chemical budget analysis are listed in Table
4. OH production by photolysis of HONO and O₃, the reaction between HO₂+NO, and between HO₂+O₃, and the limonene
ozonolysis reaction are considered. The uncertainties of these OH-producing processes range from 15% (photolysis processes)
to 20% (HO₂+NO reaction and limonene ozonolysis reaction), which are derived from error propagation of the uncertainties
595 of measurements (Table 1) and reaction rate constants (Atkinson et al., 2004; Cox et al., 2020).

Figure 8 and Figure S7 show the chemical budgets of OH radicals in the limonene oxidation experiments with different NO
mixing ratios. The destruction and production rates of OH for experiments with medium (08 August 2012) and high NO (03
August 2015) are balanced within the uncertainties of measurements during all experiments. This is consistent with the overall
good agreement between modelled and measured OH radical concentrations obtained for the constrained model run (Fig. 5
600 and 6). In the experiment with high NO, the total OH turnover rate is higher (> 20 ppbv h⁻¹) compared to the other experiments
due to the high limonene concentration (10 ppbv) and the high NO concentration, both of which accelerate the turnover of
radicals.

In experiments with medium and high NO concentrations, the main OH source after the limonene injection is the reaction
between HO₂ and NO (>85%), followed by the photolysis of HONO in the experiment with medium NO, and ozone photolysis
605 and limonene ozonolysis in the experiment with high NO, when 60ppbv O₃ was present. In total, about 85 to 100% of the OH
radical production rate can be explained by the calculated processes that are listed in Table 4.

In the experiment with low NO on 01 September 2012 (Fig. 8b), the OH production rate is 6 ppbv h⁻¹ after the injection of
limonene, of which about 33% can be attributed to OH production from limonene ozonolysis. The NO concentration varies
over the course of the experiment making OH regeneration from the HO₂+NO reaction an important OH source with a
610 contribution of 15 – 50% to the total OH production in addition to OH production from the photolysis of HONO and O₃. About
1.0 ppbv h⁻¹ (20 – 33%) of the OH destruction rate is not explained by these OH production processes, which is consistent with
the underestimation of OH concentrations in the constrained model run compared to measured values (Fig. 4). However, the
gap between the OH destruction rate and destruction rate is about 1 ppbv h⁻¹ throughout the whole experiment and does not
vary with the amount of limonene presence in the chamber. This suggests that either the missing OH source is not related to
615 the oxidation of limonene or is due to measurement artefacts for example in the OH measurements, which would lead to an
overestimation of the OH destruction rate. An artefact in the OH measurement would also be consistent with what is observed
in the constrained model run (Fig. 4), where the rapid decrease of the modelled limonene concentration suggests that the
measured OH is too high in this experiment.

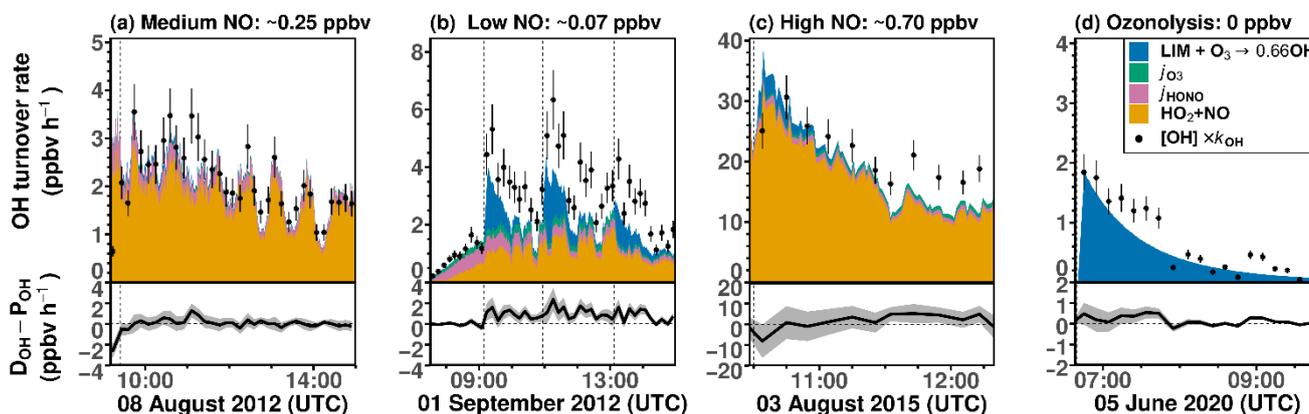
In the ozonolysis experiment without OH scavenger (Fig. 8d), OH is only produced by ozonolysis reaction of limonene. The
620 total OH production rate is about 2 to 3 ppbv h⁻¹ at the beginning of the oxidation and gradually declines while limonene is



being consumed. The total OH destruction rate is well explained by the production from limonene ozonolysis suggesting that OH production from further ozonolysis reactions of product species is not significant for conditions of this experiment.

In conclusion, the OH production rate from the four major OH sources that are included in the calculations (Table 4) is balanced by the OH destruction rate within the 25% uncertainty of the calculation at the beginning of the experiment, when limonene-oxidation is most important in the experiments with medium and high NO and in the ozonolysis experiment. In the experiments with low NO concentrations, imbalances of 20 to 33% are observed indicating that an additional OH production process with a rate of 1.0 ppbv h⁻¹ would be required to explain the observed destruction rate, but there are indications that this could be due to a measurement artefact in the OH measurements.

In all experiments, the OH production rates are lower than OH destruction rates at later times of the experiments, when secondary chemistry becomes important. However, differences are similar to the uncertainty of the calculations. These discrepancies may indicate that additional OH could be produced from unaccounted reactions of oxidation products for example from the photolysis of organic peroxides in the photo-oxidation experiments (Badali et al., 2015).



635

Figure 8. 10-minutes-average values of total OH destruction rates compared to the sum of OH production rates from the major OH sources that can be calculated from measurements in the (a) medium NO, (b) low NO, (c) high NO, and (d) ozonolysis experiments. Shaded areas in the difference plots give the uncertainties of the calculations. Production of OH from the reaction of O₃ and HO₂ is not included, because the contribution to the total OH production was negligible (< 0.01 ppbv h⁻¹) for conditions of the experiments.

640



645 **Table 4.** Reactions that are included in the analysis of the OH production rate. Reaction rate constants are given for 298K and 1 atm. The reaction rate constants that are used in the analysis are calculated for the measured temperature.

Reaction:	k (298 K, 1 atm)	1σ uncertainty of k (%)	Reference:
$\text{HO}_2 + \text{NO} \rightarrow \text{OH} + \text{NO}_2$	$8.8 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$	11	Atkinson et al. (2004)
$\text{HO}_2 + \text{O}_3 \rightarrow \text{OH} + 2\text{O}_2$	$2.0 \times 10^{-15} \text{ cm}^3 \text{ s}^{-1}$	22	Atkinson et al. (2004)
$\text{HONO} + h\nu \rightarrow \text{OH} + \text{NO}$	j_{HONO}	10	Measured
$\text{O}_3 + h\nu \rightarrow \text{O}({}^1\text{D}) + \text{O}_2$	j_{O_3}	10	Measured
$\text{O}({}^1\text{D}) + \text{H}_2\text{O} \rightarrow 2\text{OH}$	$2.1 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$	11	Atkinson et al. (2004)
$\text{O}({}^1\text{D}) + \text{M} \rightarrow \text{O}({}^3\text{P}) + \text{M}$	$2.6 \times 10^{-11} \text{ cm}^3 \text{ s}^{-1}$	11	Atkinson et al. (2004)
$\text{Limonene} + \text{O}_3 \rightarrow 0.66\text{RO}_2 + 0.66\text{OH}$	$2.1 \times 10^{-16} \text{ cm}^3 \text{ s}^{-1}$	11	Cox et al. (2020)

3.4 Chemical budget of first-generation peroxy radicals using measured quantities

The discrepancies between measured and simulated organic peroxy radical concentrations are much higher (up to a factor of
650 2) in the experiments with low NO concentrations and in the ozonolysis experiment compared to the experiments with medium or high NO concentrations. The analysis of the composition of the RO₂ concentrations using model results from the constrained model run (Section 2.4) shows that the concentrations of RO₂ produced in the initial reaction of limonene with OH and O₃ already exceed the measured total RO₂ concentrations (Fig. 9 and Supplementary Material Fig. S8). Model-measurement percentage differences are at least a factor of 2 higher than the accuracy of the measured RO₂ concentration (~25%). Therefore,
655 the discrepancy suggests that additional loss pathways for RO₂ have to be included in the model.

To examine the magnitude of the additional RO₂ loss rate, a chemical budget analysis for RO₂ radicals is performed similarly to the analysis for OH radicals (Section 3.3). As the chemical loss rates of peroxy radicals are within the range of 0.01 to 0.20 s⁻¹, steady-state concentrations can be assumed. The production rate of the peroxy radicals produced right after the limonene injection is well-defined by the loss rate of limonene due to the reactions with OH- and O₃. Therefore, only measurements
660 during the first 30 minutes after the first limonene injection are used for the analysis, so that calculations are not impacted much by additional RO₂ production from the subsequent oxidation of organic products. The removal rate of peroxy radicals includes bimolecular reactions (k_{bi}):



$$k_{\text{bi}} = k_{\text{RO}_2+\text{RO}_2}[\text{RO}_2] + k_{\text{RO}_2+\text{HO}_2}[\text{HO}_2] + k_{\text{RO}_2+\text{NO}}[\text{NO}] \quad (6)$$

665 Values of the bimolecular reaction rates in the MCM for limonene-derived RO₂ radicals are: $k_{\text{RO}_2+\text{RO}_2}$ ranging from 10⁻¹³ to 10⁻¹² cm³ s⁻¹, $k_{\text{RO}_2+\text{HO}_2} = 2.3 \times 10^{-11}$ cm³ s⁻¹ at 298K, $k_{\text{RO}_2+\text{NO}} = 9.0 \times 10^{-12}$ cm³ s⁻¹ at 298K. The additional loss rate (k_{add}) required to balance production (P_{RO_2}) and destruction (L_{RO_2}) rates can be calculated as:

$$L_{\text{RO}_2} = [\text{RO}_2](k_{\text{bi}} + k_{\text{add}}) = (0.66 \times k_{\text{lim}+\text{O}_3}[\text{O}_3] + k_{\text{lim}+\text{OH}}[\text{OH}])(\text{lim}) = P_{\text{RO}_2} \quad (7)$$

$$k_{\text{add}} = P_{\text{RO}_2}[\text{RO}_2]^{-1} - k_{\text{bi}} \quad (8)$$

670 In this calculation, measured concentrations of radicals and trace gases are used. The RO₂ yield of 0.66 from the limonene ozonolysis reaction assumes RO₂ radicals are produced from the vinyl hydroperoxide (VHP) mechanism together with an OH radical (Wang and Wang, 2021).

675 Results from the five experiments are summarized in Table 5. The values of the additional loss rate, k_{add} , range from 0.01 to 0.06 s⁻¹ in different experiments, but are similar with the high relative uncertainty of at least 50%. The large relative uncertainty is caused by the small differences between production and destruction rates, which also increase with increasing RO₂ and NO concentrations (and thereby k_{bi}). The relative uncertainties of the additional loss are less than 100% in only three experiments, which include the ozonolysis experiment (05 June 2020), the experiments with low (01 September 2012), and medium NO concentrations (08 August 2012). In the ozonolysis experiment, the additional RO₂ loss is lower by a factor of 4 than in the other 2 experiments. This might be explained by the low temperature during this experiment, which may slow down the additional loss pathway. Another explanation could be the high fraction of RO₂ from limonene ozonolysis, which may behave 680 differently from the RO₂ from the OH reaction with limonene.

Potential additional loss pathways of RO₂ radicals that are needed to balance the RO₂ production rate in the experiments would be radical conversion reactions that can also lead to the production of OH radicals, so that these pathways are potentially connected to imbalances found in the chemical budget of OH radicals (Section 3.3). It is worth noting that discrepancies in the chemical budgets of RO₂ and OH radicals are both smaller in the ozonolysis experiment compared to the photochemistry 685 experiments at low and medium NO mixing ratios.

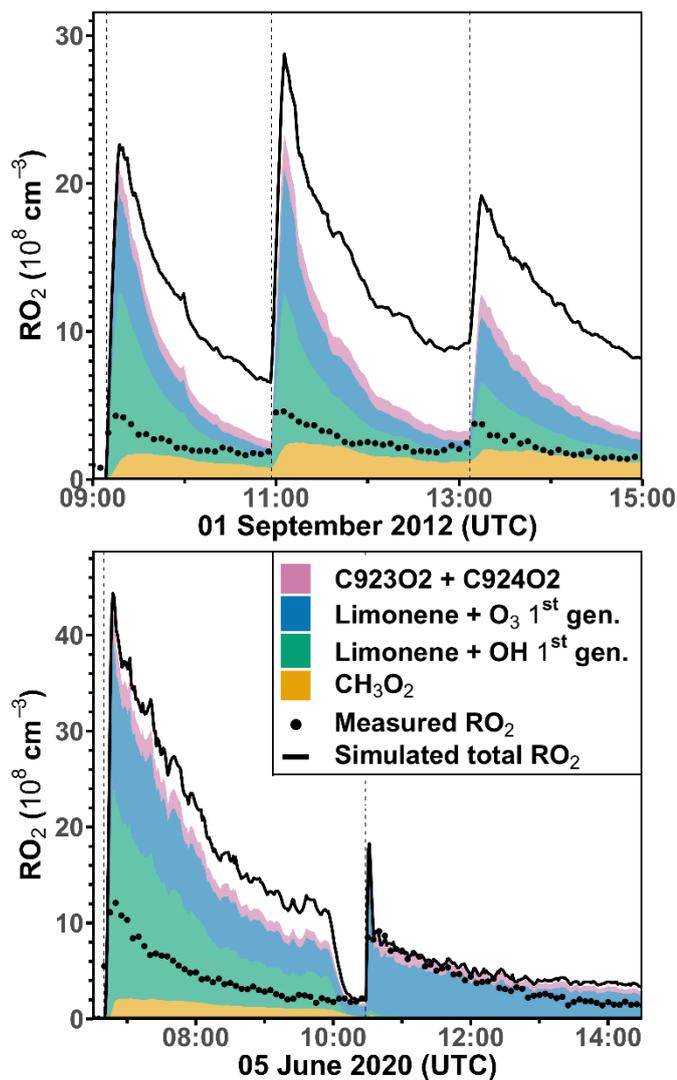


Figure 9. Total RO_2 radical concentrations and their speciation from model calculations (constrained model run) compared to the measured values (black dots) for the experiments with low NO and the ozonolysis experiment. Methylperoxy radicals (CH_3O_2) are mainly produced from the oxidation of HCHO. Radicals produced in the initial reactions of limonene with either OH or O_3 are summed. C923O2 + C924O2 are RO_2 radicals produced from the further oxidation of the first-generation oxidation products. Names are taken from the MCM model.



Table 5. Additional removal rates (k_{add}) that are required to balance the RO_2 production and destruction rates in the different experiments together with conditions of the experiments such as the percentage of limonene that reacted with OH (LIM+OH) or O_3 (LIM+ O_3). Only data from 30 minutes after the first limonene injection is analysed.

NO	Date	T (K)	LIM+OH (%)	LIM+ O_3 (%)	P_{RO_2} ($10^7 \text{ cm}^{-3} \text{ s}^{-1}$)	[RO_2] (10^8 cm^{-3})	k_{bi} (10^{-2} s^{-1})	k_{add} (10^{-2} s^{-1})
Zero	05 June 2020	286	49	51	2.3 ± 0.6	10.0 ± 1.4	0.9 ± 0.3	1.4 ± 0.7
Low	01 September 2012	313	59	41	2.8 ± 0.8	3.5 ± 0.5	2.5 ± 0.7	5.6 ± 2.7
Low	04 July 2019	302	72	28	3.1 ± 0.9	6.4 ± 1.0	3.1 ± 0.9	1.7 ± 1.7
Medium	08 August 2012	303	97	3	1.7 ± 0.5	1.4 ± 0.2	6.9 ± 2.1	5.2 ± 3.5
Medium	10 August 2012	302	95	5	1.8 ± 0.5	1.6 ± 0.2	8.2 ± 2.7	3.3 ± 4.2

3.5 Impacts of RO_2 model results on the modelled NO_x concentrations

700 In the model runs described so far, NO and NO_2 concentrations are constrained to measured values. If these are not constrained, NO_x is underestimated at the beginning of the experiments, but values are overestimated at the end (Fig. 10 and Supplementary Material Fig. S9). The discrepancy at the beginning can be mainly attributed to the overestimation of modelled RO_2 concentration, which leads to an overestimation of the formation of organic nitrates that act as sinks for NO_x on the timescale of the experiments. To illustrate the impact of RO_2 concentrations on the modelled NO_x concentrations, two model runs are
705 compared: one with modelled RO_2 concentrations (reference run) and the second with modelled RO_2 concentrations adjusted to match the measurements.

RO_2 concentrations are adjusted by applying an additional loss with a constant rate constant to all six first-generation RO_2 derived from -OH and - O_3 oxidation (Fig. 1). The additional loss rate for RO_2 is around 0.01 to 0.06 s^{-1} , similar to the loss rate derived in the analysis of the chemical budget for RO_2 (k_{add} , Section 3.4). In both model runs, the organic nitrate yield of the
710 RO_2+NO reaction for the first-generation RO_2 radicals from limonene+OH is adjusted to the yield (34%) that is found in the analysis of the experiments at medium NO mixing ratio (Section 3.1.2). Measured j_{NO_2} , O_3 and HO_2 concentrations are used in both model runs to constrain the loss rate of NO_2 and NO. It should be noted that only data within one hour after the first limonene injection is evaluated as RO_2 produced from the further oxidation of organic products are not considered.

Figure 10 shows the modelled NO and NO_x concentrations for the two model runs. In the experiment with medium NO
715 concentrations on 08 August 2012, RO_2 radical concentrations are overestimated by about 50 – 100% by the reference model and modelled NO as well as NO_x concentrations, are 25% lower than measurements. With an additional RO_2 loss rate of 0.05 s^{-1} (Table 5), the fraction of RO_2 that reacts with NO reduces from 80 – 90% to 45 – 60%. Therefore, the loss of NO_x by the formation of organic nitrates is also reduced, so that the model-measurement agreement for NO and NO_x improves for the first two hours of the experiments.



720 Both model runs overestimate the NO_x and NO concentrations, when all limonene reacted away after 13:00 UTC (Fig. 5). The measured NO_x concentration remains stable at around 0.6 ppbv throughout the whole experiment after the injection of limonene. However, NO_x concentrations increase at a rate of about 0.15 ppbv h^{-1} in the reference model. The increase is reduced to less than 0.05 ppbv h^{-1} in the model run with the additional RO_2 loss during the last two hours of the experiment. The production of NO_x in the model at later times of the experiment can be explained by the production of NO_2 from the photolysis of the first-generation organic nitrates and their oxidation by OH. These effects are more important in the reference model run, when the modelled first-generation organic nitrates are high. To reconcile the difference in NO_x concentrations between the model and measurements, a stronger nitrogen sink is required in the model. This may also suggest that the model underestimates the organic nitrate formation from the reaction of NO with RO_2 from the oxidation of product species. Another explanation would be that the lifetime of limonene-derived organic nitrates from OH-oxidation is too short in the model.

730 In the experiment with low NO concentrations on 01 September 2012, the fraction of RO_2 that reacts with NO reduces from about 50% to 16%, if an additional RO_2 loss with a rate of 0.06 s^{-1} (Table 5) is applied. In this experiment, a large fraction of NO_x in the model is lost due to the formation of PAN or PAN-like species from acyl peroxy radicals (e.g., CH_3CO_3 and $\text{C}_8\text{H}_{17}\text{CO}_3$) that are formed in the radical chain reaction of the ozonolysis reaction of limonene (Fig. 1). The additional loss of the initially formed RO_2 species competes with the reaction with NO_2 and therefore the formation of PAN reduces, if the additional loss is applied. This effect of reduced NO_x loss in the ozonolysis reaction adds to the effect for a reduced organic nitrate formation discussed for the experiment on 08 September 2012 at medium NO.

Although the reduced NO_x loss significantly improves the model-measurement agreement for the first part of the experiments, if an additional RO_2 loss process is included in the model, NO_x concentrations are overestimated by this model at later times, when the chemistry of product species gains in importance. This could be due to neglecting the impact of the subsequent chemistry of the additional RO_2 loss reactions on nitrogen oxide concentrations. The chemistry of nitrogen oxide species in the experiment with low NO concentrations is more complex compared to the experiment with medium NO as a significant fraction of RO_2 radicals is produced by the ozonolysis of limonene in addition to the reaction with OH. Further investigation will be required to specifically clarify the impact of the formation of PAN and PAN-like species from the ozonolysis of limonene (Fig. 1). To our best knowledge, there is no experimental study investigating PAN formation from the oxidation of limonene.

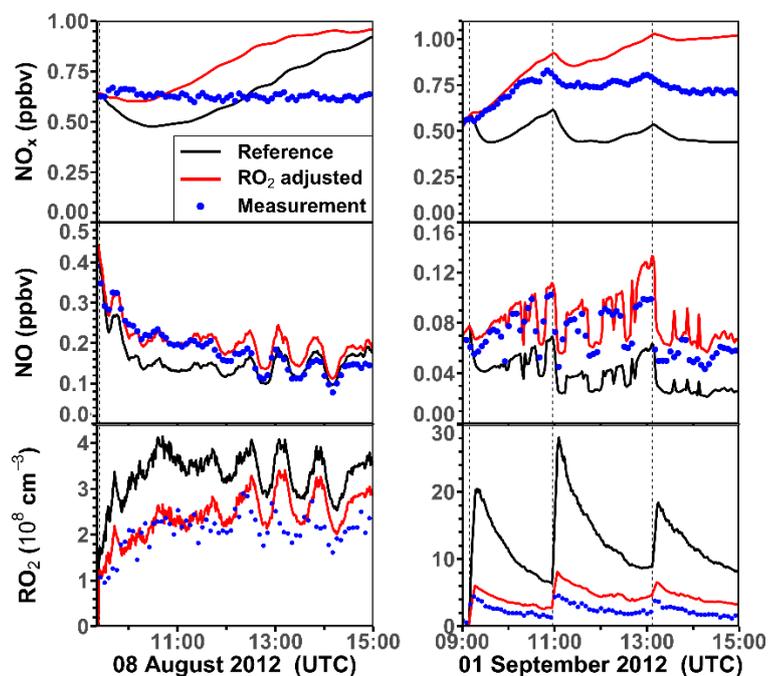


Figure 10. Example of the impact of too high modelled RO₂ on the modelled NO_x and NO concentrations in the experiments with medium NO (08 August 2012) and low NO (01 September 2012) concentrations. In both cases, the organic nitrate yield of 34% from the reaction of NO with first-generation RO₂ from the limonene+OH reaction is taken from the analysis in this work.

4 Discussion

4.1 Sensitivity model runs including additional radical regeneration reactions for RO₂ radicals

Additional radical regeneration reactions are further explored by sensitivity model runs. These model runs aim to reduce the discrepancies between measured and modelled OH, HO₂ and RO₂ concentrations in the reference model. This is achieved by implementing additional reaction pathways:



where RO₂ are the peroxy radicals distinguished from their production in the reaction of limonene with either OH (limOH-RO₂; i.e., LIMAO₂, LIMBO₂, and LIMCO₂, Fig. 1) or O₃ (limO₃-RO₂; i.e. LIMALAO₂, LIMALBO₂, and L5O₂, Fig. 1), because these peroxy radicals are structurally similar with peroxy radicals from the OH reaction having a β-OH moiety of a 6-carbon ring and limO₃-RO₂ being acyclic peroxy radicals with a β-oxo, an aldehyde and an isopropenyl group. Therefore, it



is assumed that they have similar reaction pathways. Reactions could involve an unknown reaction partner X like for example
765 used in Hofzumahaus et al. (2009) or could be unimolecular reactions.

Reaction rate constants for Reaction R1a and R1b (k_{R1a} and k_{R1b}) are implemented as pseudo-first order reaction rate constants. RO_2 within the same group (limOH- RO_2 or limO₃- RO_2) are assumed to have the same rate constants. Reaction R1a would lead to HO₂ production and Reaction R1b would lead to OH production. In the sensitivity model runs, the reaction rate constants are optimized to minimize the model-measurement discrepancies for OH, HO₂, RO_2 concentrations and OH reactivity. The
770 sum of k_{R1a} and k_{R1b} must be within the range of the additional RO_2 loss rate k_{add} (Section 4.1, Table 5). A missing HO₂ source is found in the reference model for all experiments (Fig. 4 – 7). Assuming that the loss rate of the HO₂ radical, which mainly reacts with NO, is correctly accounted for, an additional RO_2 to HO₂ conversion (Reaction R1a) is needed to bring measurement and model results in agreement. In contrast, missing OH is only found in the experiments with low NO as evident from the analysis of the chemical budget of OH radicals (Section 3.3, Fig. 8). These observations indicate that additional RO_2
775 to OH conversion (Reaction R1b) can only be competitive with other bimolecular reactions for NO mixing ratios of less than 0.05 ppbv which is equivalent to a loss rate of $k < 10^{-2} s^{-1}$.

The model-measurement agreement of radical concentrations is first optimized based on the second half of the ozonolysis experiment, when CO was added as OH scavenger. In this case, only limO₃- RO_2 are present, but the conversion to either HO₂ (Reaction R1a) or OH (Reaction R1b) cannot be distinguished, because OH rapidly converts to HO₂. To achieve agreement
780 between modelled and measured HO₂ concentrations during this part of the ozonolysis experiment, the sum of the additional loss rates ($k_{R1a} + k_{R1b}$) would need to be $(0.017 \pm 0.008) s^{-1}$. The uncertainty is mainly due to the uncertainty in the measurement of HO₂ concentrations (~20%). The upper limit for the rate constant k_{R1b} for the loss of limO₃- RO_2 can be estimated from the first part of the ozonolysis experiment, when no OH scavenger was present. Since 80 to 100% of the observed OH production can already be explained by OH production from the limonene ozonolysis reaction (Fig. 8d), the rate of k_{R1b} for limO₃- RO_2
785 would need to be less than $0.004 s^{-1}$. This implies that the rate constant k_{R1b} for limOH- RO_2 is also less than $\sim 0.004 s^{-1}$ as about 40% of limonene is oxidized by OH in the ozonolysis experiment without OH scavenger.

The implementation of the Reaction 1a for limO₃- RO_2 cannot significantly improve the model-measurement discrepancies of HO₂ concentrations in the experiments, when limonene is predominantly oxidized by OH. Also in the ozonolysis experiment, HO₂ concentrations are still underestimated by about 40% during the part of the experiment without OH scavenger. Hence,
790 the reaction rate k_{R1a} for an additional loss of limOH- RO_2 is also optimized to match the measured HO₂ concentrations.

Optimization of the reaction rate k_{R1a} for the additional loss of limOH- RO_2 for individual experiments results in values that differ by one order of magnitude. For instance, the optimum rate constant is $(0.006 \pm 0.003) s^{-1}$ in the ozonolysis experiment without OH scavenger, but it is $(0.05 \pm 0.03) s^{-1}$ in the experiment with medium NO concentrations on 08 August 2012. These
795 optimizes rate constants are consistent with the values of the loss rate k_{add} (Table 5), with the rate required in the ozonolysis experiment having a slower rate while the rate required in the experiment with medium NO having a faster rate. The exact



reason for such large differences is not clear, but could be related to the higher temperature (16 – 27°C) in the photo-oxidation experiments, when the chamber air was exposed to sunlight. The average value of the rate constant for the conversion from RO₂ to HO₂ for the experiments in this work is 0.03 s⁻¹. This value is applied to all sensitivity model runs in the following to illustrate its impacts on modelled RO₂, HO₂ concentrations and OH reactivity (Table 6).

800 A summary of all reactions included for the sensitivity run is available in Table 6. Figure 11 shows the increase in the OH production rate in the sensitivity model runs that include the conversion of limO₃-RO₂ to OH at a rate of 0.004 s⁻¹. The total OH production rate increases by about 0.2 ppbv h⁻¹ in both experiments corresponding to a 5% and 10% increase, respectively. This reduces the imbalance between OH production and destruction rates in the experiment with low NO by about 20% without significantly impacting the balance in the ozonolysis experiment. This demonstrates that the additional OH production from
805 the conversion of first-generation RO₂ from OH- or O₃- oxidation of limonene to OH is not sufficient to fully close the gap between OH production and destruction rates in the experiment with low NO, for which the discrepancy is largest among all experiments in this work.

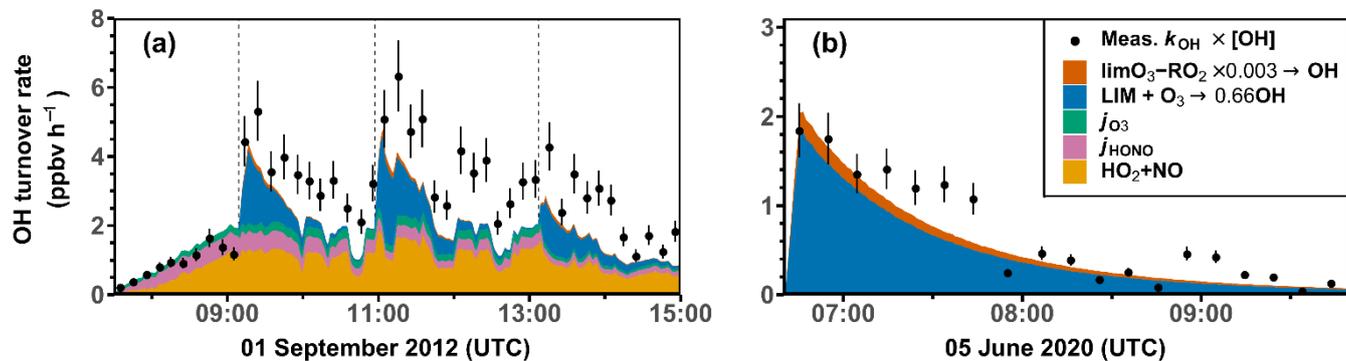
Figure 12 and Figure S10 show radical concentrations and OH reactivity obtained in the reference and sensitivity model runs. In the sensitivity model run, the model-measurement agreement for RO₂ and HO₂ concentrations improves compared to the
810 reference model run as can be expected from the adjustment of the reaction rate constant. In the experiment with low NO concentrations, however, an optimal agreement of both, RO₂ and HO₂ concentrations, cannot be simultaneously achieved. This suggests that some fraction of the additional RO₂ loss pathway may not regenerate HO₂ or OH radicals.

In the sensitivity model run the overestimation of the OH reactivity is reduced even without introducing additional loss pathways of oxidised products as implemented in the constrained model run (Section 2.4), because the production of organic
815 peroxides is reduced due to the competition with the additional RO₂ loss reaction. For example, the percentage of RO₂ reacting with HO₂ reduces from about 50% to 25% and from 90% to 30% in the experiment with low NO and in the ozonolysis experiment, respectively. In the sensitivity model, no closed-shell products are produced from the additional RO₂ loss reaction. The good model-measurement agreement of the OH reactivity suggests that organic products from these reactions are not reactive or they are rapidly lost for example to the chamber wall. Therefore, no further conclusions, about the type of products
820 formed from Reaction R1 can be drawn from these experiments.



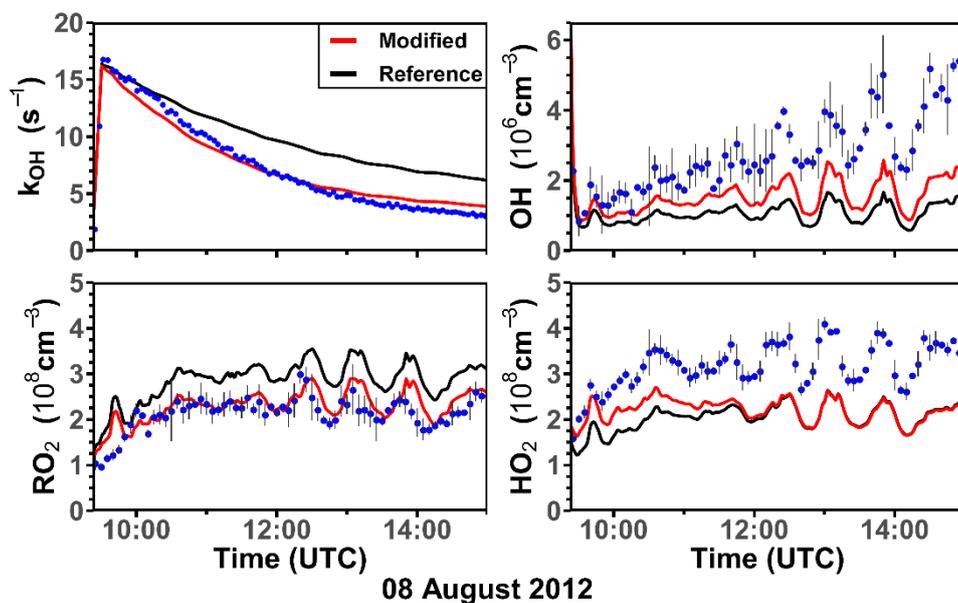
Table 6. Modification of chemical reactions implemented in the sensitivity simulations runs.

Reaction	Reaction rate constant	Comment
LIMALAO2 → OH	0.003 s ⁻¹	Illustrate the impact of the additional OH source on the OH budget in the ozonolysis experiment, when there is no OH scavenger.
LIMALBO2 → OH	0.003 s ⁻¹	
L5O2 → OH	0.003 s ⁻¹	
LIMALAO2 → HO2	0.014 s ⁻¹	Derived from the optimisation of HO ₂ model-measurement agreement in the ozonolysis experiment when there is OH scavenger. Assuming k_{R1b} for limO ₃ -RO ₂ is 0.003 s ⁻¹
LIMALBO2 → HO2	0.014 s ⁻¹	
L5O2 → HO2	0.014 s ⁻¹	
LIMAO2 → HO2	0.030 s ⁻¹	Mean value of the rate derived from the optimisation of HO ₂ model-measurement agreement in the experiments with low NO, medium NO and the ozonolysis experiment.
LIMBO2 → HO2	0.030 s ⁻¹	
LIMCO2 → HO2	0.030 s ⁻¹	

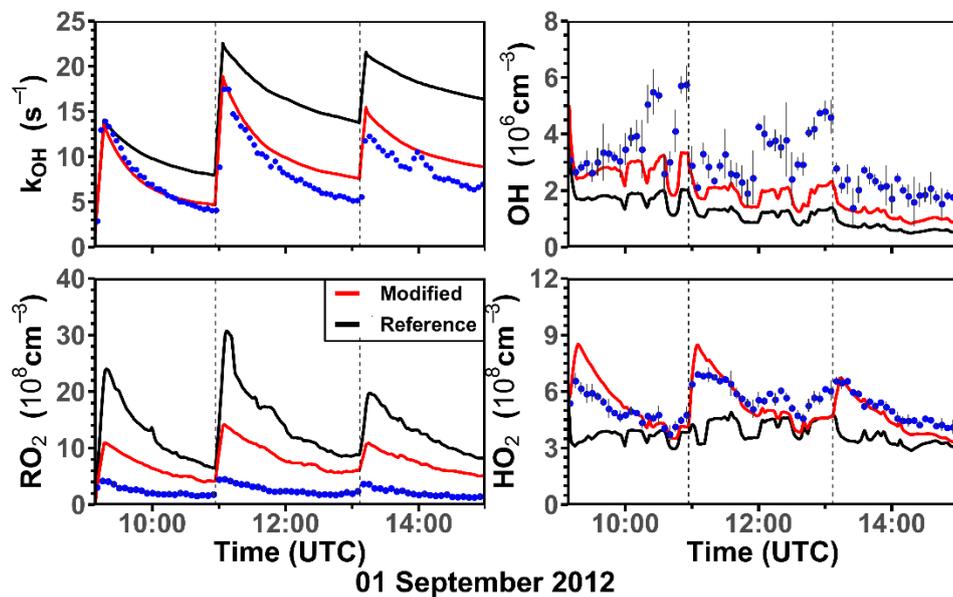


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Figure 11. Measured 10-minutes-mean total OH production rate compared with the OH production rate from the main measured OH sources for (a) the experiment with low NO concentration on 01 September 2012 and (b) the ozonolysis experiment on 05 June 2020 for the sensitivity model run that includes additional OH production from the reaction of RO₂ from limonene ozonolysis and HO₂ (Reaction R2c).



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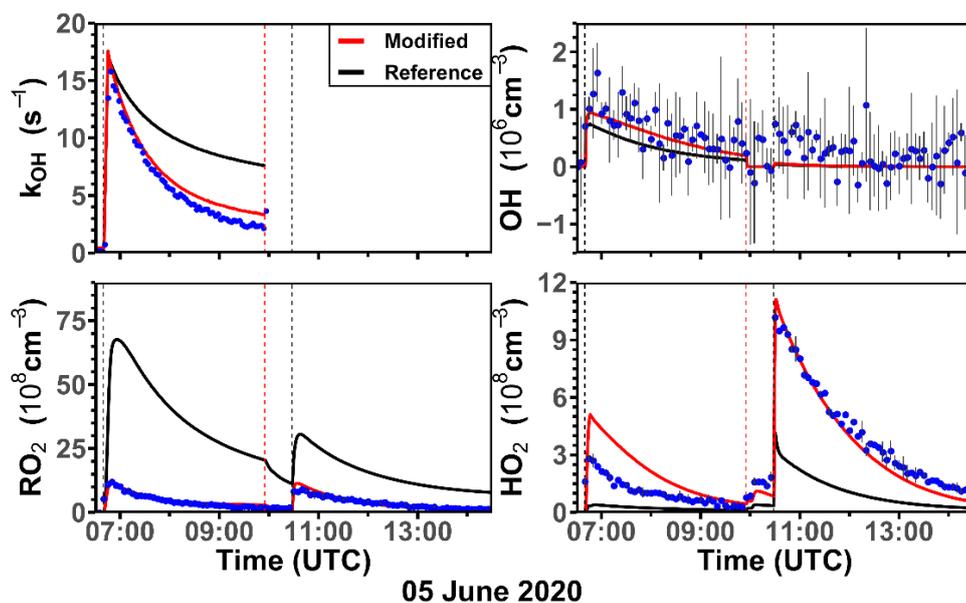


Figure 12. Modelled and measured OH reactivity, HO₂, RO₂, and OH concentrations for the experiments with medium NO (top) and low NO (middle) concentrations, and for the ozonolysis experiment (bottom). Results from the reference model (reference) and the sensitivity model run (modified) that includes additional RO₂ loss processes producing OH and HO₂ (Table 6) are compared.

4.2 Possible additional OH sources from RO₂ reactions

Possible underlying mechanisms of the additional RO₂ loss and the production of OH are discussed in the following section. It should be noted that the additional OH source discussed here is referring to the slow additional OH source ($k_{R1b} < 0.003 \text{ s}^{-1}$) that could only slightly reduce the discrepancy between the OH production and destruction rates (Fig. 11). The larger discrepancy in the chemical OH budget that is observed in the experiment with low NO concentrations (Fig. 11a) compared to the other experiments cannot be explained by the additional conversion of RO₂ to OH and, hence, it is not further discussed here. The rate of the additional RO₂ to OH conversion, k_{R1b} , is much slower than the total loss rate of RO₂ due to bi-molecular reactions ($k_{bi} > 0.01 \text{ s}^{-1}$) or the additional RO₂ to HO₂ conversion ($k_{R1a} \sim 0.017 \text{ s}^{-1}$). Its contribution to the production rate of OH would also be small in all experiments ($< 10\%$ of the total OH production rate). Therefore, there is no strong evidence for an additional conversion of RO₂ to OH for experimental conditions in this study.



4.2.1 RO₂ isomerisation reactions

855 Isomerisation of RO₂ could lead to the production of OH, which is shown in the OH-oxidation of methacrolein (e.g., Crouse et al., 2012) and isoprene (e.g., Novelli et al., 2020). Isomerisation reactions for RO₂ from the limonene oxidation by OH and O₃ are investigated by Møller et al. (2020) and Chen et al. (2021), respectively. Of the peroxy radical LIMAO₂, LIMBO₂ and LIMCO₂ that are produced from the limonene oxidation by OH, it was calculated that LIMAO₂ and LIMBO₂ could undergo -OH H-shift with a rate constant of 10⁻³ to 10⁻⁴ s⁻¹. LIMCO₂ could undergo isomerisation reactions that are competitive with
860 bimolecular reactions for atmospheric conditions, which include a cyclization reaction with a rate constant of 0.2 to 0.8 s⁻¹ and an allylic-H shift reaction with a rate constant of 0.1 to 1.7 s⁻¹ (Fig. 13). On the other hand, peroxy radicals LIMALAO₂, LIMALBO₂, and L5O₂ that are produced from the ozonolysis of limonene could all undergo much faster isomerisation reactions ($k > 0.5$ s⁻¹) such as aldehydic H-shift and cyclisation reactions (Fig. 14) due to the loss of steric hinderance after the ring-opening ozonolysis reaction.

865 Subsequent reaction steps after the first isomerisation reaction were not investigated in the work by Møller et al. (2020) and by Chen et al. (2021). Therefore, potential OH production from subsequent reactions are estimated with available SAR (e.g., Vereecken and Peeters, 2009; Vereecken and Nozière, 2020; Vereecken et al., 2021).

The subsequent reaction of β-OOH substituted alkoxy radicals, LIMA_15shift_O and LIMB_15shift_O, which are products of the 1,5 -OH H-shift reactions of LIMAO₂ and LIMBO₂ respectively, is a ring-breaking decomposition reaction that
870 produces an OH radical. For the cyclization product from the isomerisation of LIMCO₂, LIMC_6cyc_O₂, it may either undergo bimolecular reactions or isomerisation by abstracting the α-H atom of the hydroxyl group (Fig. 13). Isomerisation of LIMC_6cyc_O₂ could result in an aldehyde together with an HO₂ radical. However, it is uncertain, if the isomerisation reaction of LIMC_6cyc_O₂ can compete with other reaction channels, as the SAR does not apply to bicyclic compounds. The reaction between LIMC_6cyc_O₂ and NO results in a bicyclic alkoxy radical (LIMC_6cyc_O), which then dissociates to a cyclic
875 intermediate, analogous to the bicyclic alkoxy radical produced from the OH-oxidation of β-pinene (Vereecken and Peeters, 2004). The dissociation of LIMC_6cyc_O may eventually produce an HO₂ radical, which could not explain the production of OH at zero NO condition.

For limO₃-RO₂, their subsequent reactions after the first isomerisation reaction are also expected to be very fast ($k > 0.1$ s⁻¹) because of the presence of either an aldehyde group or a C=C double bond. However, allylic H-shift, cyclisation, and aldehydic
880 H-shift reactions do not affect the RO₂ and OH concentrations as the resulting alkyl radicals do not decompose into a closed-shell product and an OH radical. Instead, an oxygen molecule rapidly adds to the alkyl radical site to form a highly oxidised RO₂.

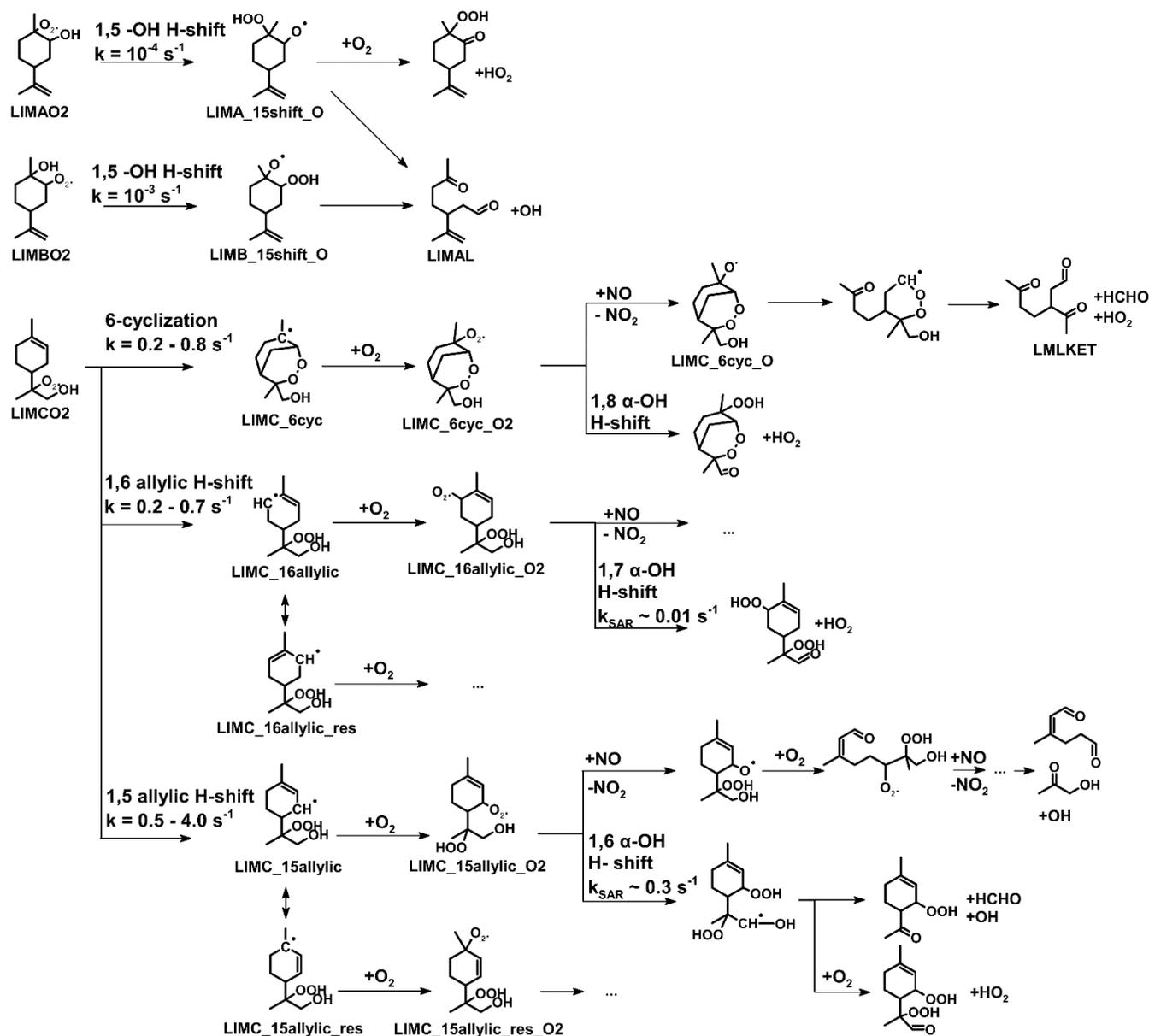
One of the possible additional OH sources from RO₂ isomerisation reactions could be from the H-abstraction of the α-OOH group (Bianchi et al., 2019). In this case, the α-OOH substituted alkyl radical could quickly decompose into a carbonyl product
885 and an OH radical. Because of the fast first and second step of the isomerisation reaction, it is expected that limO₃-RO₂ could



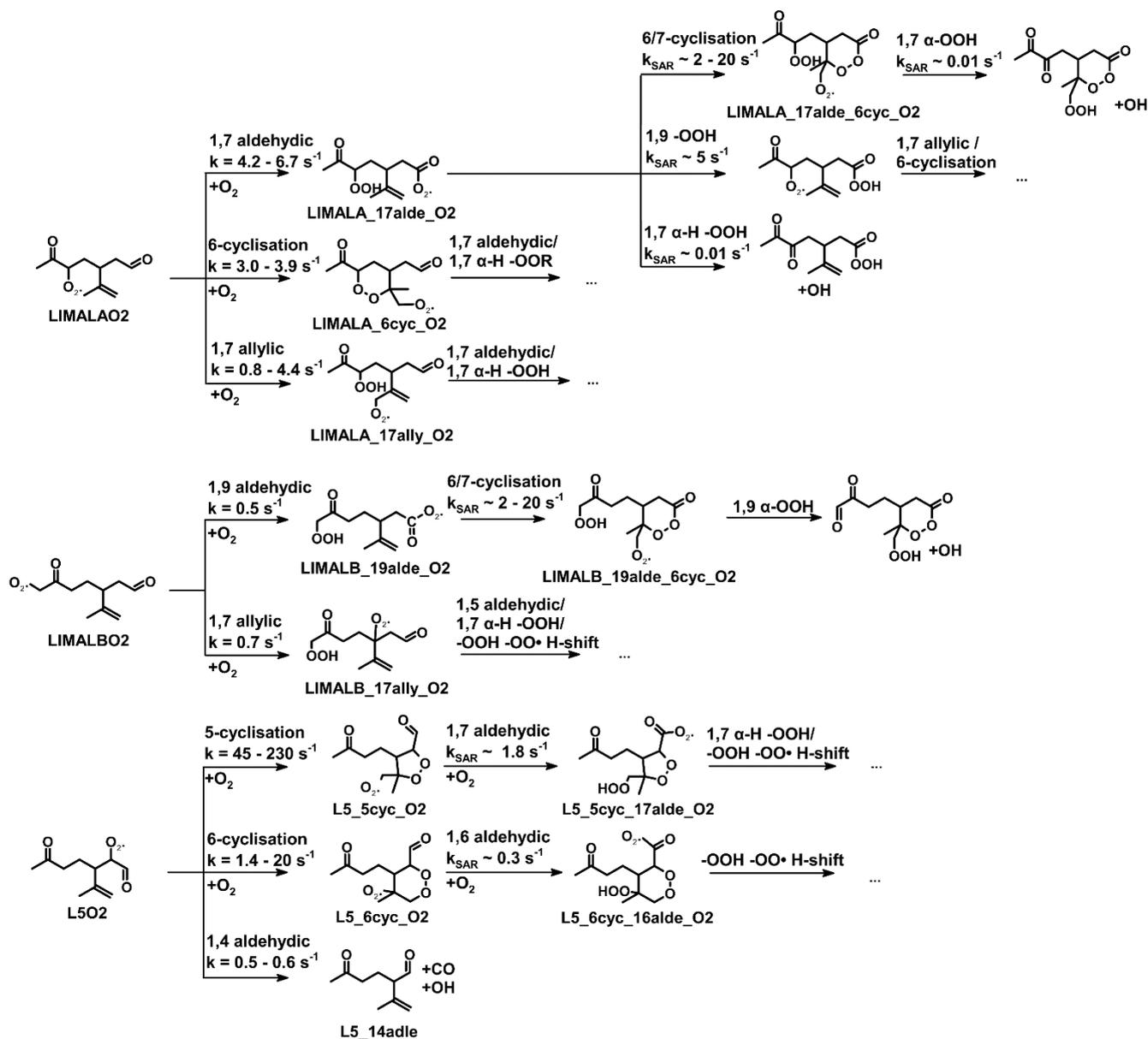
form RO₂ with a hydroperoxide group that allows them to undergo α-OOH H-shift reaction (e.g., LIMALA_17alde_O2, LIMALB_19alde_O2, L5_5cyc_17alde_O2 (Fig. 14)). The SAR-estimated α-OOH H-abstraction rate is about 10⁻³ to 10⁻² s⁻¹ (Vereecken and Nozière, 2020), which is slightly faster than the additional RO₂ loss rate k_{Rib} (< 0.003 s⁻¹) applied in the sensitivity run. It should be noted that the SAR does not consider the impacts of neighbouring functional groups on the α-OOH H-shift rate. All limO₃-RO₂ have a β-oxo group near the radical site, which could increase the H-abstraction rate near the β-oxo group by about one order of magnitude (e.g., Crouse et al., 2013), including the α-OOH H-abstraction. However, RO₂ derived from the isomerisation of limO₃-RO₂ still have other possible isomerisation reaction pathways that are even more competitive ($k > 1$ s⁻¹) than the α-OOH H-shift. One examples is the H-shift reaction from the hydroperoxyl group to another peroxy group (Vereecken and Nozière 2020). This complicates the chemistry of these highly oxidized peroxy radicals and it is unclear, whether they eventually undergo an α-OOH H-abstraction.

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905



910 **Figure 13.** Unimolecular reaction pathways of RO₂ radicals from the reaction of limonene with OH reported in Møller et al. (2020) and subsequent anticipated reaction pathways. Reaction rate constants denoted with k_{SAR} are directly taken from SAR for peroxy radicals (Vereecken and Nozière, 2020). The unimolecular reaction rate for 1,8 α -OH H-shift is not available.



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Figure 14. Unimolecular reaction pathways for the major limonene-ozonolysis-derived RO₂ radicals. Rates taken from calculations in Chen et al. (2021), in which stereoisomers are considered and therefore a range of values are given. Other rates (*k*_{SAR}) are based on SAR in Vereecken and Nozière (2020) and Vereecken et al. (2021). It should be noted that their SAR does not address the effect of steric hindrance (in the case of a cyclic compound) or multiple functional groups near the abstracted H-atom on the H-shift rate constant.

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4.2.2 Reactions between RO₂ and HO₂

Another possible RO₂ loss reaction that leads to OH production and is not considered in the reference model is the reaction of β-oxo-substituted RO₂ with HO₂. Based on the work by Jenkin et al. (2019), RO₂+HO₂ reactions can lead to different products depending on the functional groups nearby the peroxy group of the RO₂ radical:



930 It has been shown in experimental studies investigating simple β-oxo substituted RO₂ that they can form an alkoxy radical together with an OH radical (Reaction R2c) instead of a hydroperoxide compound (Reaction 2a) (Jenkin et al., 2007; Hasson et al., 2012; Praske et al., 2015). With the branching ratios that are taken from the SAR, which is derived using simple β-oxo substituted RO₂, the OH yield of the RO₂+HO₂ reaction for limO₃-RO₂ is about 30% (Jenkin et al., 2019). This value is similar to the OH production rate in the sensitivity runs using an additional RO₂ loss rate of $k_{\text{R1b}} = 0.003 \text{ s}^{-1}$. However, theoretical
935 investigation by Iyer et al. (2018) suggests that the OH yield of the RO₂+HO₂ reaction may not be high enough (< 1%) based on the energy barriers that were calculated in that study for limO₃-RO₂. Currently, there is no laboratory study on the OH yield from the RO₂+HO₂ reaction for large RO₂ and therefore the OH yield of the RO₂+HO₂ for large RO₂ is highly uncertain.

4.3 Possible additional HO₂ sources from RO₂ reactions

940 4.3.1 RO₂ isomerisation reactions

Similar to the additional OH source, HO₂ can also be produced from the isomerisation reaction of RO₂ radicals (e.g., Crouse et al., 2012; Peeters et al., 2014). Again, the possibility of additional conversion from RO₂ to HO₂ by isomerisation is investigated using the isomerisation pathways calculated for RO₂ derived from limonene oxidation and SARs (Vereecken and Peeters, 2009; Møller et al., 2020; Vereecken and Nozière, 2020; Chen et al., 2021; Vereecken et al., 2021).

945 Peroxy radicals LMAO₂ and LIMBO₂ that are produced from the OH-oxidation of limonene can undergo a slow ($k < 10^{-3} \text{ s}^{-1}$) -OH H-shift reaction that is one order of magnitude slower than the rate of the RO₂ to HO₂ conversion applied in the sensitivity run (Fig. 13). In addition, the production of HO₂ through the H-abstraction by O₂ of the alkoxy radical LIMA_15shift_O is not as favourable as the ring-cleavage alkoxy dissociation that eventually produces an OH radical



(Vereecken and Peeters, 2009). Therefore, even with the potential production of HO₂ through the isomerisation of RO₂ derived
950 from the isomerisation of LIMCO₂ (e.g., 1,8 α-OH H-shift of LIMC_6cyc_O₂), the production of HO₂ is limited by the 37%
yield of LIMCO₂ from the oxidation of limonene by OH. In addition, the rate constant of the H-shift reaction for the
intermediate radicals derived from LIMCO₂ (e.g., LIMC_6cyc_O₂, LIMC_16allylic_O₂, LIMC_15allylic_O₂) have a high
uncertainty as it is assumed that the SAR for acyclic compounds can be applied for cyclic RO₂. For these reasons, isomerisation
reactions for limOH-RO₂ are unlikely the reason for the additional HO₂ production and RO₂ loss required to match observed
955 radical concentration measurements.

For limO₃-RO₂, reaction rate constants of the first two of the isomerisation reaction are about one to two orders of magnitude
faster than the RO₂ to HO₂ conversion rate used in the sensitivity run (Fig. 14). As discussed in Section 4.2.1, one of the
possible RO₂ loss mechanisms through isomerisation reaction is a α-OOH H-abstraction reaction. Although the value of the
rate constant of the α-OOH H-abstraction reaction derived from SAR is on the same magnitude (~ 10⁻² s⁻¹) as the RO₂ to HO₂
960 conversion rate applied in the sensitivity run, abstraction of the hydrogen with an α-OOH group would lead to the production
of an OH radical rather than a HO₂ radical. Therefore, isomerisation reactions of limO₃-RO₂ can also not explain the missing
RO₂ to HO₂ conversion resulting from observations in the experiments in this work.

4.3.2 Reaction rate of the RO₂ recombination reaction

965 Apart from isomerisation, HO₂ could also be produced from the dissociation of alkoxy radicals derived from RO₂ from the
reaction of limonene with OH. Alkoxy radicals could be produced from the recombination reaction of RO₂ radicals in addition
to the reaction of RO₂ with NO:



The current knowledge about the branching ratio between Reaction R3a to R3d, as well as the RO₂ self- or cross-reaction rate
constants is limited especially for complex RO₂ derived from monoterpenes. There are no specific investigations for RO₂ from
limonene. Reaction rate constants implemented in the MCM model are based on estimated cross-reaction rates between RO₂
975 and methyl peroxy radicals (CH₃O₂), (Jenkin et al., 1997, 2019). The reaction rate constants of the RO₂ recombination
reactions, $k_{\text{RO}_2+\text{RO}_2}$, for limonene-derived radicals are between 10⁻¹² to 10⁻¹³ s⁻¹ in the MCM consistent with results for RO₂
from methyl cyclohexene, which contain a tri-substituted endocyclic double bond like limonene (Boyd et al., 2003).



980 However, the reaction rate constant $k_{\text{RO}_2+\text{RO}_2}$ could be higher, if the cross-reaction partners are other large limonene-derived radicals rather than CH_3O_2 . For example, Berndt et al., (2018) investigated the self-reaction rate constants for RO_2 derived from the reaction of α -pinene with OH after they undergo two steps of unimolecular reactions (i.e., $\text{C}_{10}\text{H}_{16}\text{OH}(\text{O}_2)_2\text{-O}_2$). They found that values range between 1 and $4 \times 10^{-11} \text{ cm}^3 \text{ s}^{-1}$ in this case. However, it should be noted that these high rates are derived from the production rate of peroxide products (ROOR, Reaction R3d), rather than the loss rate of RO_2 .

985 Using the values of the reaction rate constants $k_{\text{RO}_2+\text{RO}_2}$ for RO_2 from limonene oxidation from the MCM, the upper limit of the RO_2 loss rate due to $\text{RO}_2\text{-RO}_2$ reactions is about 10^{-3} s^{-1} in the ozonolysis experiment and experiments with low NO, and $2 \times 10^{-4} \text{ s}^{-1}$ in the experiments with medium NO. From the additional loss rate ($\sim 10^{-2} \text{ s}^{-1}$; Table 5) determined from the chemical budget analysis for RO_2 , the value of the reaction rate constant for $\text{RO}_2\text{-RO}_2$ reaction that would be required to explain the observations ($k'_{\text{RO}_2+\text{RO}_2}$) can be calculated. This results in values of $k'_{\text{RO}_2+\text{RO}_2}$ that are about 3×10^{-10} , 1×10^{-11} , $3 \times 10^{-11} \text{ cm}^3 \text{ s}^{-1}$ in the medium NO, low NO and ozonolysis experiments, respectively. The uncertainties of the rates are about 50 – 60%, which are derived from the error propagation of RO_2 concentrations and optimal rate constants in the sensitivity model run (Table 6).
990 It should be noted that these values are collective loss rates of all first-generation RO_2 species from limonene oxidation before the formation of closed-shell products, including highly-oxidized RO_2 produced from potential auto-oxidation reactions.

The values of the reaction rate constant $k'_{\text{RO}_2+\text{RO}_2}$ found in the low NO experiment and ozonolysis experiment are in the same order of magnitude (10^{-11} to $10^{-10} \text{ cm}^3 \text{ s}^{-1}$) as values reported by Berndt et al., (2018) for RO_2 from α -pinene oxidation. Berndt et al., (2018) also showed that the reaction rate constant for the $\text{RO}_2\text{-RO}_2$ self-reaction increases, when the RO_2 becomes more
995 oxidised. This hints the importance of RO_2 recombination reactions for RO_2 derived from limonene oxidation could be higher than previously thought, because of the rapid isomerisation reaction of these radicals (Møller et al. (2020) and Chen et al. (2021); Section 4.2.1).

In the experiment with low NO concentrations, the additional loss rate for RO_2 radicals that is required to explain measured RO_2 concentrations ($k \sim 0.06 \text{ s}^{-1}$, Table 5) is higher than the rate of the additional RO_2 to HO_2 conversion required to explain
1000 measured HO_2 concentration ($k \sim 0.006 - 0.02 \text{ s}^{-1}$). This would be consistent with a faster reaction rate constant for the $\text{RO}_2\text{-RO}_2$ reaction, because only a fraction of $\text{RO}_2\text{-RO}_2$ reaction would lead to the formation of alkoxy and therefore HO_2 radicals (Reaction 3).

However, it would be unclear, why the reaction constant $k'_{\text{RO}_2+\text{RO}_2}$ required in the experiment with medium NO mixing ratio would be higher compared to other experiments. It is also worth noting that the decomposition of alkoxy radicals produced
1005 from RO_2 from the ozonolysis of limonene leads to the production of peroxy radicals which do not lead to the production of HO_2 in most of the $\text{RO}_2\text{-RO}_2$ reaction chain (Fig. 1). Therefore the missing production of HO_2 in the ozonolysis experiment cannot be explained by a higher than previously thought reaction rate constant of the $\text{RO}_2\text{-RO}_2$ reaction.



5 Conclusions

1010 The photooxidation of limonene by OH and O₃ was investigated in experiments for zero, low (~0.1 ppbv), medium (~0.4 ppbv) and high (~ 1 ppbv) NO levels in the atmospheric simulation chamber SAPHIR. The experiments were conducted with limonene mixing ratios of 4 – 10 ppbv and O₃ mixing ratios ranging from 0 – 50 ppbv.

The analysis of measured radical concentrations in the experiments revealed that current knowledge about the limonene oxidation as implemented in the Master Chemical Mechanism cannot explain observed values specifically concerning radical
1015 regeneration. Observed OH and HO₂ concentrations were a factor of 2 to 3 higher than predicted by model calculations, whereas measured RO₂ concentrations were at least 50% lower than modelled values. The following processes in the limonene mechanism impacting radical concentrations could be identified that are not appropriately described:

- The loss rate of OH radicals is too high in the model as seen in higher-than-observed OH reactivity values in the model. Although it cannot be excluded that chamber wall losses reduced the concentration of organic oxidation
1020 products, this hints that the reactivity of products species with respect to their reaction with OH is low or other products other than currently thought are produced for example by the competition of unaccounted radical reaction pathways.
- The yield of organic nitrates from the reaction of RO₂ radicals formed in the initial reaction of OH with limonene is found to be (34±5)%, which agrees with the measurements by Rollins et al. (2010), but which is about 10% higher
1025 than calculated from structure-activity relationships. The higher yield of organic nitrates reduces the efficiency of the radical regeneration in the limonene mechanism.
- Formaldehyde is expected to be formed from the reaction chain after the addition of OH to the terminal C=C double bond, if RO₂ radicals react with NO, so that the formaldehyde yield would be similar to the yield of that RO₂ species (37%). The low formaldehyde yield of (13±3)% in the experiments with medium NO concentrations suggests that
1030 there is an unaccounted RO₂ loss reaction not producing formaldehyde that is competitive at 200 pptv NO.
- OH production and destruction rates are balanced in most of the experiments, if measured OH reactivity and measured HO₂ concentrations are used for calculating reaction rates. This demonstrates that measured values are consistent and confirm the shortcomings of the limonene mechanism with regards to describing the HO₂ production and OH reactivity.
- An unaccounted RO₂ loss process with a rate of 0.02 to 0.06 s⁻¹ is required to balance the RO₂ production rate from
1035 the reaction of OH with organic compounds. Formation of HO₂ with a rate of 0.03 s⁻¹ and 0.017 s⁻¹ from an additional reaction of the RO₂ from the reaction of limonene with OH and O₃, respectively, can explain part of the model-measurement discrepancies for HO₂.
- An unaccounted RO₂ loss process for RO₂ from the ozonolysis of limonene that is competitive against the reaction
1040 with NO prevents the formation of NO_x reservoir species PAN and PAN-like species as suggested in the MCM model.



The observed NO_x concentrations do not exhibit a distinct temporal behaviour that would be expected neither from the rapid loss of NO_x species at the beginning of limonene oxidation when RO_2 derived from limonene oxidation reacts with NO_x , and NO_x reformation from the thermal decomposition of PAN species at the later times of the experiments when limonene has reacted away.

1045 Overall, the results of the experiments clearly demonstrate that loss reactions of RO_2 from the oxidation of limonene are not well understood. Unaccounted RO_2 reactions lead to an enhanced radical regeneration. Organic products likely are less reactive than products that are currently thought to be formed. Time series of measured radical concentrations indicate that their further oxidation reactions need to be investigated to explain observed values at later times of the experiments, when limonene had reacted away. The formation of NO_x reservoir species (PAN / PAN-like species) is lower than expected. However, this is partly
1050 counter-acted by a high yield of $(34\pm 5)\%$ of organic nitrates. Oxidation products are also likely to have a high organic nitrates yield as indicated by the low measured NO_x concentrations that would otherwise be expected to continuously increase over the course of the experiment due to the continuous emission of HONO by the chamber film.

Rates of RO_2 isomerisation reactions proposed by Møller et al. (2020) and Chen et al. (2021) are too low or too fast ($k \sim 10^{-4}$ to 10^{-3} and 1 to 10^2 s^{-1}) to explain observed radical concentrations and expected products are not consistent. A possible
1055 explanation could be that the reaction rate constant $k_{\text{RO}_2+\text{RO}_2}$ of RO_2 recombination reactions for RO_2 from limonene oxidation is higher than the reaction rate constants that are calculated by SAR implemented in the MCM. For example, experiments for RO_2 from α -pinene oxidation by Berndt et al. (2018) show that the rate constant can be one to two orders of magnitude faster than implemented in the MCM model for α -pinene (10^{-13} to $10^{-12} \text{ cm}^3 \text{ s}^{-1}$). These values are consistent with the additional loss rate required to explain radical concentrations in the experiments with limonene in this work. However, the importance of the
1060 alkoxy pathway of the RO_2+RO_2 reaction for large monoterpene-derived RO_2 is still unclear and needs further investigation.

Data availability: Data from the experiments in the SAPHIR chamber used in this work are available on the EUROCHAMP data home page. 08 August 2012: <https://doi.org/10.25326/JNMN-YC22> (Fuchs et al., 2021b); 10 August 2012: <https://doi.org/10.25326/2PGS-FP66> (Fuchs et al., 2021c); 01 September 2012: <https://doi.org/10.25326/77N2-ZK22> (Fuchs et al., 2021a); 03 August 2015: <https://doi.org/10.25326/BP56-WP95> (Bohn et al., 2021c), 04 July 2019: <https://doi.org/10.25326/C4SW-TP73> (Bohn et al., 2021b); 05 June 2020: <https://doi.org/10.25326/7ZKX-347> (Bohn et al., 2021a).

Supplement. The supplement related to this article is available online at:

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Author contributions. JYSP, AN and HF wrote the manuscript. MK, AN and HF designed and led the experiments in the chamber. BB (radiation), RT, AL, IHA and RW (organic compounds), XL (HONO), FR (ozone, nitrogen oxides), HPD and PC (radicals), SN (OH reactivity), CC and AN (radicals and OH reactivity) were responsible for measurements used in this work. All co-authors commented and discussed the manuscript and contributed to the writing of the manuscript.

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