Author Response for "Aqueous SOA formation from the photo-oxidation of vanillin: Direct photosensitized reactions and nitrate-mediated reactions" by Mabato et al.

We thank the reviewer for the thorough review and many constructive comments that helped improve the manuscript. Our point-by-point responses are below (changes to the original manuscript text and supporting information are in red, moved content in double-line strikethrough, and removed content in strikethrough). Please note that the line numbers in the responses refer to our revised manuscript with tracked changes. Also, please note that because we restructured the manuscript, the numbering of some figures and tables in the revised manuscript is different from those in the original manuscript.

Reviewer 1

Overview

The authors examined the aqueous photodegradation of vanillin (VL), a carbonyl-containing phenol emitted from biomass burning, and accompanying formation of aqueous SOA (aqSOA). They then measured the composition of the aqSOA using high-resolution mass spectrometry and UV/Vis absorption. They also determined the impact of purging solutions with N2 (to remove dissolved oxygen) as well as the addition of ammonium nitrate (a photochemical source of hydroxyl radical, OH) and/or one of two hydroxyl radical scavengers (isopropyl alcohol or bicarbonate). They attempt to explain their results qualitatively based on a few dozen reactions, but there is little experimental attempt to text the mechanisms.

Vanillin has been studied in several past works, but this paper adds new information on the composition of the resulting aqSOA. The purging with N2 is novel, but the interpretation of the results is not clear, and I disagree that these experiments show that secondary oxidants dominate VL loss. Unfortunately, the nitrate concentration added was too low to impact kinetics (because VL direct photodegradation is so fast), but it's interesting that it impacted the products formed. Finally, the authors seem compelled to try to mechanistically explain most of their results, but their explanations are very speculative and should be significantly cut. There are a several other major and minor issues, as described below.

Response: In this study, we aimed to investigate the photo-oxidation of VL at atmospherically relevant cloud and fog conditions. As mentioned in lines 112–116, 'Although the concentration of VL in cloud/fog water has been estimated to be < 0.01 mM (Anastasio et al., 1996), a higher VL concentration (0.1 mM) was used in this study to guarantee sufficient signals for product identification (Vione et al., 2019). The chosen ammonium nitrate (AN) concentration (1 mM) was based on values observed in cloud and fog water (Munger et al., 1983; Collett et al., 1998; Zhang and Anastasio, 2003; Li et al., 2011; Giulianelli et al., 2014; Bianco et al., 2020).' Our study is not intended to identify the concentrations of nitrate that would affect the kinetics. This sentence has been added to the text as follows:

Line 116: It should be noted that this study is not intended to identify the concentrations of nitrate that would affect the kinetics.

Major Comments

1. The normalized abundance of products (line 131) is used throughout the paper as a key metric, but it's unclear if this is a robust endpoint, in part because its uncertainty is never discussed. (a) Based on the major products that have been identified (both via MS and IC), what is the likely range of ionization efficiencies (IEs) of the products and how much uncertainty does this introduce in the product abundance measure? (b) There is additional uncertainty in the quantification of VL, which is described on line 134 as semi-quantitative. (c) Altogether, what is the relative uncertainty in P from day to day and experiment to experiment? (d) This is an issue because there are several times when the normalized product abundance results are inconsistent with other, seemingly more quantitative metrics. For example, in section 3.1.3., the presence of OH scavengers had no significant effect upon VL decay or aqSOA light absorbance, but there were differences in the normalized abundance of products. Given the uncertainty in IE and other aspects of the product measure, I would be wary of attributing much significance to the normalized abundance of products as an endpoint when it's inconsistent with the more quantitative measures.

Response: The normalized abundance of products in this study is a semi-quantitative analysis intended to provide an overview of how the signal intensities (as normalized in Eq. 2) changed under different experimental conditions, but not to quantify the absolute concentration of products. Even if relative abundance (product peaks are normalized to the highest peak), which has been widely utilized (e.g., Lee et al., 2014, https://pubs.acs.org/doi/10.1021/es502515r; Romonosky et al., 2017, https://pubs.acs.org/doi/10.1021/acs.jpca.6b10900; Fleming et al., 2018, https://doi.org/10.5194/acp-18-2461-2018; Klodt et al., 2019, https://pubs.acs.org/doi/pdf/10.1021/acsearthspacechem.9b00222) in the literature is used instead of normalized abundance in our analysis, the major products detected as well as the conclusions of this study will remain the same. Regardless, we agree with the reviewer that we should emphasize the potential uncertainties in the normalized abundance of products. Detailed responses to the relevant sub-questions are as follows:

(a) Ionization efficiencies can indeed vary between different compounds. Unfortunately, the availability of measured relative ionization efficiencies (RIE) for different compounds is limited. We are not in a position to provide this information. The reviewer is correct that ESI ionization is not ideal for product quantification. Nevertheless, Nguyen et al. (2013) (https://doi.org/10.1039/C2AY25682G) found a positive correlation between ESI signal and "adjusted mass" (= molecular mass × H: C). Based on that study, the uncertainty would be a factor of 2 – 4 if only the "adjusted mass" is considered, and further complications of matrix effect and polarity are disregarded. However, what we compared is not the absolute concentrations (or contributions) of the products observed. The comparison was based on how the signal intensities (as normalized in Eq. 2) changed under different experimental conditions. We compared the

responses of the same products (or at least the same class of products) as conditions varied, and their ionization efficiency might not be very different within the same class, according to the "adjusted mass" concept by Nguyen et al. (2013). We have revised Sect. 2.2 to highlight the inherent uncertainties for this metric due to ionization efficiencies which can vary for different compounds as follows:

Lines 146–167: Comparisons of peak abundance in mass spectrometry have been used in many recent studies (e.g., Lee et al., 2014; Romonosky et al., 2017; Wang et al., 2017; Fleming et al., 2018; Song et al., 2018; Klodt et al., 2019; Ning et al., 2019) to show the relative importance of different types of compounds (Wang et al., 2021). However, ionization efficiency may greatly vary for different classes of compounds (Kebarle, 2000; Schmidt et al., 2006; Leito et al., 2008; Perry et al., 2008; Kruve et al., 2014) and so uncertainties may arise from comparisons of peak areas among compounds. In this work, we assumed equal ionization efficiency of different compounds, which is commonly used to estimate O:C ratios of SOA (Bateman et al., 2012; Lin et al., 2012; De Haan et al., 2019), to calculate their normalized abundance. The normalized abundance of a product, [P] (unitless), was calculated as follows:

$$[\mathrm{P}] = \frac{\mathrm{A}_{P,t}}{\mathrm{A}_{VL,t}} \cdot \frac{[\mathrm{VL}]_t}{[\mathrm{VL}]_0} \qquad \text{(Eq. 2)}$$

where $A_{P,t}$, and $A_{VL,t}$ are the extracted ion chromatogram (EIC) ~~signal~~ peak areas of the product P and VL from UHPLC-qToF-MS analyses at time t, respectively; $[VL]_t$ and $[VL]_0$ are the VL concentrations (μM) determined using UHPLC at time t and 0, respectively. Here, we relied on the direct quantification ~~more accurate measurements~~ of [VL] using UHPLC (see Fig. S2 for VL calibration curve) ~~for semi-quantification~~. ~~It should be noted that the ionization efficiency may greatly vary for different classes of compounds (Kebarle, 2000). Hence, we assumed equal ionization efficiency of different compounds to calculate their normalized abundance, which is commonly used to estimate O:C ratios of SOA (Bateman et al., 2012; Lin et al., 2012; De Haan et al., 2019).~~ It should be noted that the normalized abundance of products in this study is a semi-quantitative analysis intended to provide an overview of how the signal intensities changed under different experimental conditions but not to quantify the absolute concentration of products. Moreover, the major products detected in this study are probably those with high concentration or high ionization efficiency in the positive ESI mode. The use of relative abundance (product peaks are normalized to the highest peak) (e.g., Lee et al., 2014; Romonosky et al., 2017; Fleming et al., 2018; Klodt et al., 2019) would yield the same major products reported. Typical fragmentation behavior observed in MS/MS spectra for individual functional groups from Holčapek et al. (2010) are outlined in Table S1.

(b) The semi-quantification here refers to the normalized abundance of products, not the VL concentration. [VL] was directly quantified using UHPLC (see Fig. S2 for VL calibration curve). The reported [VL] are the average of results from triplicate experiments and the uncertainties from which and those from the MS signal intensities were propagated (now added in Table 2, formerly S2). This has been clarified in the text as follows:

Line 157: Here, we relied on the direct quantification ~~more accurate measurements~~ of [VL] using UHPLC (see Fig. S2 for VL calibration curve) ~~for semi-quantification~~.

(c) Given the same instrumental settings, the variations caused by the instrumental fluctuations would be smaller than the effects caused by the difference in ionization efficiency among different species. If there is any, it would be taken into account by the normalization in Eq. 2, which is why the normalized signal intensities were used instead of absolute signal intensities. For reference, relative uncertainties for MS signal peak areas of VL at the same concentration measured from different experiments range from 0.14 to 0.25. Moreover, the propagated uncertainties from the MS signal intensities and [VL] are now shown in Table 2 (formerly S2).

(d) For $^{\bullet}$OH scavengers experiments, the insignificant changes for VL decay and absorbance enhancement might not be reflected in the products observed using UHPLC-qToF-MS in positive ESI mode. It is possible that the products observed might not have contributed significantly to all products formed and may not be the primary contributors to the absorbance enhancement. The absorbance enhancement may not necessarily correlate directly with the products detected. However, as mentioned in our response to major comment #4, we decided to omit the section for $^{\bullet}$OH scavengers based on the likely minor contribution of $^{\bullet}$OH to VL photo-oxidation in this study.

2. Throughout the manuscript, the low decay rate of VL* under N2 is taken to mean that the triplet state of VL isn't involved in VL decay and that secondary, O2-dependent, oxidants are responsible for VL decay. However, the N2-purging control experiment result is ambiguous, since secondary steps in VL decay via triplets might require oxygen to proceed. For example, a major fate of the ketyl radical formed by the 3VL* + VL reaction is to add oxygen. In the absence of oxygen, the ketyl radical will still form, but it's forward path (O2 addition) is blocked, possibly leading to eventual return to the reactants (and little apparent VL decay). So N2 purging is likely to not only remove secondary oxidants, but also to interfere with subsequent steps in the 3VL* - VL reactions. Thus the oft-stated conclusion that secondary oxidants from 3VL* are responsible for VL decay is not correct (e.g., on line 184). Without knowing the impact of O2 on the reaction intermediates in the triplet reaction, it is impossible to know what the N2-purging result means.

Another strike against the "secondary oxidants" theory is that the proposed secondary oxidants are unlikely to be important for VL decay. For example, the 1O2* + VL reaction is slow under the pH conditions here (where there is negligible phenolate). In air-saturated solutions, the 1O2* and 3VL* concentrations should be roughly equal (see the McNeil and Canonica review in ESPI), but at pH 4 (and below) the rate constants for phenols with 3C* are much faster than the 1O2* values. The bottom line is the 1O2* is unlikely to be important. Similarly, HO2/O2- was proposed as an oxidant for phenols, but these are very weak oxidants that react slowly with phenols. Finally, OH is apparently unimportant as well, based the OH scavengers having no significant impact on VL decay; however, it is possible that most of the IPA or bicarbonate was

purged from the sample prior to illumination (as discussed below). Regardless, photolysis of H2O2 (formed from the 3VL* + VL reaction) will be slow, giving little OH.

Response: Thank you for this thorough and important analysis. We apologize for the confusion related to the role of VL triplets. The reviewer is correct that VL triplets are indeed important for VL decay and that the secondary oxidants generated in the presence of O_2 likely have only minor roles in the photo-oxidation of VL in this study. We also agree that without a detailed investigation of the effect of O_2 on the reactive intermediates, it is difficult to interpret the mechanism of the N_2 experiments. In principle, initial oxidation by triplets can proceed without O_2, forming phenoxy (which is in resonance with a carbon-centered cyclohexadienyl radical that has a longer lifetime) and ketyl radicals (Neumann et al., 1986a, 1986b; Anastasio et al., 1996). The coupling of phenoxy radicals or phenoxy and cyclohexadienyl radicals can form oligomers, as observed for both N_2- and air-saturated experiments. However, the little decay of VL under N_2-saturated condition indicates that these radicals probably predominantly decayed via back-hydrogen transfer to regenerate VL (Lathioor et al., 1999). A possible explanation for this is the involvement of O_2 in the secondary steps of VL decay, likely concerning the fate of the ketyl radical, as the reviewer pointed out. We have amended the discussions to include these possibilities and to emphasize the importance of VL triplets as follows:

Line 187: ~~Effect of secondary oxidants from VL~~ VL photo-oxidation under N_2 and air-saturated conditions

Lines 188–220: ~~As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions), while ^{3}VL* is the only oxidant expected under N_2-saturated conditions.~~ The photo-oxidation of VL ~~To examine the contributions of ^{3}VL*-derived secondary oxidants and ^{3}VL* only on VL photo-oxidation, experiments~~ under both N_2-~~air-~~ and air-~~N_2-~~saturated conditions (Fig. S3a) were carried out at pH 4, which is representative of moderately acidic aerosol and cloud pH values (Pye et al., 2020). No significant VL loss was observed for dark experiments. The oxidation of ground-state VL by ^{3}VL* via H-atom abstraction or electron transfer can form phenoxy (which is in resonance with a carbon-centered cyclohexadienyl radical that has a longer lifetime) and ketyl radicals (Neumann et al., 1986a, 1986b; Anastasio et al., 1996). The coupling of phenoxy radicals or phenoxy and cyclohexadienyl radicals can form oligomers as observed for both N_2- and air-saturated experiments (see discussions later). However, the little decay of VL under N_2-saturated condition indicates that these radicals probably predominantly decayed via back-hydrogen transfer to regenerate VL (Lathioor et al., 1999). A possible explanation for this is the involvement of O_2 in the secondary steps of VL decay. For instance, a major fate of the ketyl radical is reaction with O_2 (Anastasio et al., 1996). In the absence of O_2, radical formation occurs, but the forward reaction of ketyl radical and O_2 is blocked, leading to the regeneration of VL as suggested by the minimal VL decay. Aside from potential inhibition of secondary oxidants generation (Chen et al., 2020), N_2 purging may have also hindered the secondary steps for VL decay.

~~The low decay rate for VL* under N_2-saturated conditions suggests a minimal role for ^{3}VL* in VL photo-oxidation.~~ Contrastingly, the VL* decay rate constant under air-saturated conditions was 4 times higher~~, revealing the importance of ^{3}VL* derived secondary oxidants for photosensitized oxidation of VL~~. As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions). However, the photo-oxidation of VL in this study is likely mainly governed by ^{3}VL* and that these secondary oxidants have only minor participation. ~~Aside from $^{\bullet}OH$, $O_2^{\bullet-}$/$^{\bullet}HO_2$ and $^1O_2$ can also promote VL photo-oxidation (Kaur and Anastasio, 2018; Chen et al., 2020).~~ $^1O_2$ is also a potential oxidant for phenols (Herrmann et al., 2010; Minella et al., 2011; Smith et al., 2014), but $^1O_2$ reacts much faster (by ~60 times) with phenolate ions compared to neutral phenols (Tratnyek and Hoigne, 1991; Canonica et al., 1995; McNally et al., 2005). Under the pH values (pH 2.5 to 4) considered in this study, the amount of phenolate ion is negligible, so the reaction between VL and $^1O_2$ should be slow. Interestingly, however, $^1O_2$ has been shown to be important in the photo-oxidation of 4-ethylguaiacol (pK_a = 10.3) by ^{3}C* of 3,4-dimethoxybenzaldehyde (solution with pH of ~3) (Chen et al., 2020). Furthermore, while the irradiation of other phenolic compounds can produce H_2O_2, a precursor for $^{\bullet}OH$ (Anastasio et al., 1996), the amount of H_2O_2 is small. Based on this, only trace amounts of H_2O_2 were likely generated from VL* (Li et al., 2014) under-air saturated conditions, suggesting that contribution from $^{\bullet}OH$ was minor. Overall, these suggest that VL photo-oxidation in this study is driven by ^{3}VL*. Further study on the impact of O_2 on the reactive intermediates involved is required to understand the exact mechanisms occurring under air-saturated conditions. Nonetheless, the VL* decay trends clearly indicate that O_2 is important for efficient VL photo-oxidation ~~an efficient oxidant for unsaturated organic compounds and has a lifetime that is much longer than ^{3}C* (Chen et al., 2020)~~.

Revisions made elsewhere in the text:

Line 23: The effects of oxygen (O_2) ~~secondary oxidants from ^{3}VL*~~

Line 25: Our findings show that ~~the secondary oxidants (^{1}O2, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) from the reactions of ^{3}VL* and~~ O_2 plays an essential role in VL photo-oxidation.

Line 91: The influences of O_2 ~~secondary oxidants from VL triplets,~~

Line 184: In this work, the direct photosensitized oxidation of VL (~~by ^{3}VL* or secondary oxidants from ^{3}VL* and O_2~~VL only experiments)

Lines 237–240: Although we have no quantification of the oxidants in our reaction systems as it is outside the scope of this study, these observations clearly substantiate that ~~secondary oxidant from ^{3}VL*, which are formed when O_2 is present, are required for efficient~~ photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation are more efficient in the presence of O_2.

Lines 247–254: ~~Compared to N_2-saturated conditions, the normalized abundance of products such as o~~Oligomers, ~~and~~ functionalized monomers (e.g., demethylated VL; Fig. S4), and nitrogen-

containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3~~3, Table S3~~, Table S2) (for VL+AN) had higher normalized abundance ~~were also more relatively abundant~~ under air-saturated conditions ~~were significantly higher under air-saturated conditions~~ (Figs. 1c-d), likely due to efficient ~~the secondary oxidants from~~ ^{3}VL*-initiated oxidation and enhanced VL nitration in the presence of ~~and~~ O_2 ~~and their interactions with nitrate photolysis products~~. ~~The nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S3) were also more relatively abundant under air-saturated conditions.~~ For both VL* and VL+AN under air-saturated conditions, the most abundant product was $C_{10}H_{10}O_5$ (No. 4, Table S2~~3~~), a substituted VL.

Lines 287–289: Among experiments A5 to A8~~ (Table S2)~~, VL+AN under air-saturated conditions (A7) had the highest normalized abundance of products and <OS_c>, ~~most~~ probably due to the combined influence of ~~the secondary oxidants from~~ ^{3}VL* and enhanced VL nitration in the presence of O_2~~, and nitrate photolysis products~~.

Lines 295–300: In brief, the presence of ~~secondary oxidants from ^{3}VL* and~~ O_2 increased the normalized abundance of products and promoted the formation of more oxidized aqSOA. ~~These trends were reinforced in the presence of nitrate, indicating synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ Compared to N_2-saturated condition, the higher normalized abundance of nitrogen-containing products under air-saturated condition for VL+AN (at pH 4) suggests a potential enhancement of VL nitration in the presence of O_2.

Lines 324–326: Overall, these trends establish that ~~secondary oxidants from ^{3}VL* and~~ O_2 ~~are~~ is necessary for the efficient formation of light-absorbing compounds from both VL* and VL+AN.

Lines 493–495: Aside from the potential imidazole derivative ($C_5H_5N_3O_2$; No. 5~~10~~, Table S2~~3~~), $C_8H_9NO_3$ (No. 2, Table S2) was also observed from VL+AN but only under N_2-saturated conditions (Fig. 1b), probably due to further oxidation by ~~secondary oxidants from~~ ^{3}VL*.

Line 512: This enhanced GUA decay rate constant may be due to the ~~following main reactions:~~ oxidation of GUA by ^{3}VL* ~~(or the secondary oxidants it generates upon reaction with O_2),~~

Lines 520: which may be due to competition between ground-state VL and GUA for reactions with ^{3}VL* ~~(or the secondary oxidants it generates upon reaction with O_2)~~

Line 543: possibly due to both GUA and ground-state VL being available as oxidizable substrates for ^{3}VL* ~~and the secondary oxidants it can generate~~.

Line 601: Our results indicate that the photo-oxidation of VL is influenced by O_2 ~~secondary oxidants from VL triplets~~, pH

Lines 602–612: ~~Compared to~~ Under N_2-saturated conditions, the absence of O_2 likely hindered the secondary steps in VL decay (e.g., reaction of ketyl radical and O_2), regenerating VL as suggested by the minimal VL decay ~~more efficient VL photo-oxidation was observed under air-~~

~~saturated conditions (O_2 is present), which can be attributed to the generation of secondary oxidants (e.g., $^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) from $^3VL^*$. Further enhancement of VL photo-oxidation under air-saturated conditions in the presence of nitrate indicates synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ In contrast, $^3VL^*$-initiated reactions proceeded rapidly under air-saturated conditions (O_2 is present) as indicated by higher VL decay rate constant and increased normalized abundance of products. For pH 4 experiments, the presence of both O_2 and nitrate resulted in the highest normalized abundance of products (including N-containing compounds) and $<OS_c>$, which may be due to O_2 promoting VL nitration. Nevertheless, further work is necessary to assess the effect of O_2 on the reactive intermediates involved in $^3VL^*$-driven photo-oxidation and elucidate the mechanisms of VL photo-oxidation under air-saturated conditions.

3. Mechanism discussion. The authors seem compelled to try to explain all of their observations using one or more reactions, but since there is no quantitative examination of these mechanisms, they are all very speculative and mostly not useful. Worse, in some (many?) cases, the proposed mechanisms are inconsistent with some of the data. Fundamentally, without building a kinetic model of the mechanism and testing it against the observations, it is difficult to know whether the proposed reactions are important. The authors put too much emphasis on trying to mechanistically explain their observations and these explanations end up being mostly conjectures that are not grounded in data. These mechanistic speculations should be greatly reduced, especially if they are inconsistent with the kinetic or light absorption data and/or if they rely primarily upon the "normalized abundance of products" metric, which seems highly uncertain.

For example, on line 226, what trends were reinforced in the presence of nitrate? Looking at Table S2, ammonium nitrate has no effect on the kinetics, does not change the normalized product abundance at pH 2.5 or 3 (but does increase it at pH 4), and has no impact on OS(C). Later, in Fig. 2, we see that the presence of nitrate only negligibly increased the long-wavelength absorbance of the products. Overall, the bulk of the observations suggest that nitrate has a minor impact on VL decay, consistent with the fast direct photodegradation of VL.

Response: We concur with the reviewer that the current manuscript contains several speculative reaction mechanisms. As building a kinetic model of the mechanisms is beyond the scope of this study, we proposed major pathways for aqSOA formation instead (Fig. 2, formerly 4). As suggested in major comment #7, Fig. 2 has been shown for the first time when potential aqSOA formation pathways were discussed, then referred to throughout the text.

The trends in line 297 (formerly 226) pertain to nitrate enhancing the increased normalized abundance of products and formation of more oxidized aqSOA from VL photo-oxidation in the presence of O_2 (VL* and VL+AN under air-saturated conditions) at pH 4, suggesting a potential enhancement of VL nitration in the presence of O_2. This has been revised as follows:

Lines 295–300: In brief, the presence of ~~secondary oxidants from $^3VL^*$ and~~ O_2 increased the normalized abundance of products and promoted the formation of more oxidized aqSOA. ~~These~~

~~trends were reinforced in the presence of nitrate, indicating synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ Compared to N_2-saturated condition, the higher normalized abundance of nitrogen-containing products under air-saturated condition for VL+AN (at pH 4) suggests a potential enhancement of VL nitration in the presence of O_2.

4. I am concerned that the authors purged IPA and bicarbonate from solution during each experiment since solutions were bubbled continuously. Do they have any way to know if these OH scavengers were removed before or during illumination? Similarly, guaiacol shouldn't undergo direct photochemical loss under illumination above 300 nm, so the apparent decay measured in the dark could be evaporation during purging. If the purging was slow enough, each bubble would achieve Henry's law equilibrium with the solution, which would allow you to estimate the rates of IPA and bicarbonate (lost as CO2) from the rate constant for GUA loss and the ratio of Henry's law constants for GUA and OH scavenger. For GUA, which can be measured by HPLC, the authors should report the fraction of the initial concentration (0.1 mM) that was lost after the 30 min of purging in the dark and the fraction then lost in the dark control for the illumination experiment. Then for IPA and bicarbonate, some estimate of their fraction lost during purging would be helpful. At the very least, this issue needs to be raised and addressed.

Response: The reviewer has correctly pointed out that the contribution of $^{\bullet}OH$ to VL photo-oxidation in this study is likely minor, which is also suggested by other published literature (Anastasio et al., 1996; Li et al., 2014). We, therefore, decided to omit this section (and related sentences) and instead focus on the other findings of the paper.

5. Section 3.1.2. (a) Are VL (and GUA) decay rate constants normalized for photon fluxes? (b) Given the variability in kinetic decays, are the relative small differences in decay rate constants between pH 2.5 and 4 statistically different? (c) Do the authors have a good measure of the variability of the kinetics, e.g., the standard deviation of j(VL) based on triplicate experiments? Given that the decays are not first order, it is more difficult to discern differences in rate constants, so I would be cautious.

Response: (a) We thank the reviewer for pointing this out. The decay rate constants were initially not normalized for photon fluxes, although the reported values are the average of results from triplicate experiments. The values have been updated to the photon flux-normalized decay rate constants, and the following information were added to the text:

Line 142: The decay rate constants were normalized to the photon flux measured for each experiment through dividing *k'* by the measured 2-nitrobenzaldehyde (2NB) decay rate constant, *j*(2NB) (see Text S6 for more details).

(b) Yes, the differences in the decay rate constants between pH 2.5 and 4 (VL*: decay rate constant at pH 2.5 is 1.6 times higher than at pH 4; VL+AN: decay rate constant at pH 2.5 is 1.4 times higher than at pH 4) are small but statistically significant ($p < 0.05$). In addition, the variability for the decay rate constant measurements among triplicate experiments for all

conditions in this study is low (the standard deviation for each condition is now added in Table 2, formerly S2). We have added this in the text as follows:

Line 331: The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$).

Other relevant changes in the text are as follows:

Line 510: The enhancement of GUA decay rate constant in the presence of VL is statistically significant ($p < 0.05$), while that in the presence of AN is not ($p > 0.05$).

(c) Yes, the reported decay rate constants are the average of results from triplicate experiments, and the standard deviation for each condition is now added in Table 2, formerly S2. The footnote of Table 2 has been revised as follows:

Table 2: [b]The data fitting was performed in the initial linear region. Each value is the average of results from triplicate experiments. Errors represent one standard deviation.

6. Lines 283-289. I would be surprised if deprotonation of phenols is responsible for the higher absorbance of the aqSOA at pH 4 compared to pH 3 and 2.5. For one, the pKa values of methoxy-substituted phenols are near 10, so there's no appreciable phenolate at pH 4. Nitro-substituted phenols can have much lower pKas, but absorbance of the aqSOA formed in the presence of nitrate is nearly the same as in the absence of nitrate, so it seems nitrophenols are a minor part of the light absorption. Another possibility is that different products are made at pH 2.5 compared to pH 4. Measuring the pH dependence of the aqSOA formed at pH 2.5 and 4 would allow you to determine whether the pH dependence is rooted in acid-base chemistry of the products or of the reactions.

Response: Thank you for the suggestion. As suggested, to understand the pH effect further, we measured the pH dependence of the aqSOA formed from VL* at pH 4 and 2.5. The reviewer is correct that deprotonation of phenols does not sufficiently explain the higher absorbance enhancement observed at pH 4 compared to pH 2.5. Based on the comparable pH dependence of the aqSOA formed from VL* at pH 4 and 2.5 (see figure below), the pH dependence observed is likely due to the acid-base chemistry of the reactions, probably involving ^{3}VL* or the excimer of VL (Smith et al., 2016). Smith et al. (2016) reported that the direct photodegradation rate constants for 0.005 mM VL at pH ≤ 3 are nearly two times lower than at pH ≥ 5. The opposite trend observed in this study for 0.1 mM VL (VL* decay rate constant at pH 2.5 is 1.6 times higher than at pH 4) may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). Also, it has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). Changes on the text are as follows:

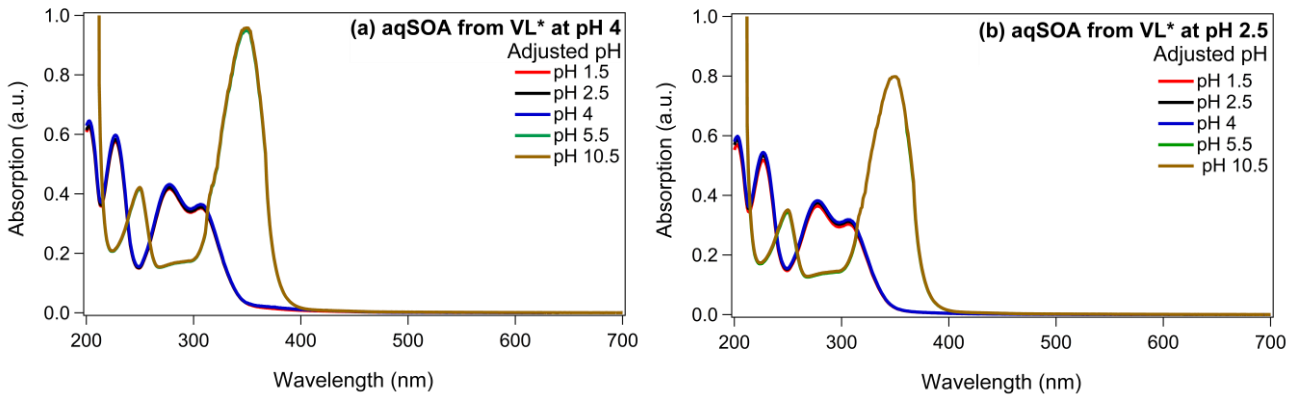

Figure S10. UV-Vis absorption spectra of VL*-derived aqSOA formed at (a) pH 4 and (b) pH 2.5 over a range of pH conditions from 1.5 to 10.5.

Lines 375–386: ~~The h~~Higher absorbance enhancement for both VL* and VL+AN (Fig. 3~~2~~b) was observed as pH increased ~~may be attributed to redshifts and increased visible light absorption of reaction products (Pang et al., 2019a~~). To determine whether the pH dependence is due to the acid-base chemistry of the products or of the reactions, we measured the pH dependence of the aqSOA formed from VL* at pH 4 and 2.5 over a range of pH conditions from 1.5 to 10.5 (Fig. S10). For both cases, the intensity of absorption at longer wavelengths significantly increased as the pH of the solutions was raised. Moreover, the comparable pH dependence of the two solutions suggest that the observed pH dependence may be attributed to the acid-base chemistry of the reactions, which may involve ^{3}VL* or the excimer of VL (Smith et al., 2016), as discussed earlier. ~~When a phenolic molecule deprotonates at higher pH, an ortho- or para- electron-withdrawing group, such as a nitro or aldehyde group, can attract a portion of the negative charge towards its oxygen atoms through induced and conjugated effects, leading to the extension of chromophore from the electron-donating group (e.g., -O$^-$) to the electron-withdrawing group via the aromatic ring (Carey, 2000; Williams and Fleming, 2008; Pang et al., 2019a). Hence, the delocalization of the negative charge in phenolates leads to significant redshifts (Mohr et al., 2013).~~

For reference, changes in the revised text that are relevant to lines 375–386 are shown below:

Lines 331–352: The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$). The pK_a for the VL triplet has been reported to be 4.0 (Smith et al., 2016). As there are a greater fraction of VL triplets that are protonated at pH 2.5 (0.96) than at pH 4 (0.5), it is possible that the pH dependence of the decay rate constants observed in this study is due to ^{3}VL* being more reactive in its protonated form. Smith et al. (2016) also observed a pH dependence for the direct photodegradation of VL (0.005 mM) (rate constants at pH ≤ 3 are ~two times lower than at pH ≥ 5) which they attributed to the sensitivity of the excimer of VL (i.e., the charge-transfer complex

formed between an excited state VL molecule and a separate ground state VL molecule; Birks, 1973, Turro et al., 2010) to acid-base chemistry. The opposite trend observed in this study for 0.1 mM VL may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). ~~For VL*, this pH trend indicates that ^{3}VL* are more reactive in their protonated form, which is opposite to that reported for 0.005 mM VL (Smith et al., 2016), likely due to the concentration dependence of the relative reactivities of protonated and neutral forms of ^{3}VL*.~~ It has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). ~~Also, increases in hydrogen ion concentration can enhance the formation of $HO_2^{\bullet}$ and H_2O_2 and in turn, $^{\bullet}$OH formation (Du et al., 2011). In addition to these pH influences on VL*, the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a) also contributed to the observed pH effects for VL+AN. At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}$OH formation than that of NO_2^- in the near-UV region (Arakaki et al., 1999; Kim et al., 2014). The increased $^{\bullet}$OH formation rates as pH decreases can lead to faster VL decay (Pang et al., 2019a). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}$OH (Reactions 4 and 15; Table 1) (Pang et al., 2019a).~~ As pH decreases, the higher reactivity of ^{3}VL* and sensitivity of the excimer of VL to acid-base chemistry ~~HONO being the dominant N(III) species can lead to faster VL photo-oxidation~~ may have led to faster VL photo-oxidation.

7. Section 3.5. This section repeats what has been stated before. I would delete this section, show Figure 4 the first time discussing possible mechanisms, then refer to the Figure throughout the discussion of mechanisms (which is hopefully much shorter in the revised version).

Response: We agree with the reviewer. Section 3.5 has been deleted, and Fig. 4, now 2, has been shown for the first time when potential aqSOA formation pathways were discussed, then referred to throughout the text as follows:

Line 259: This compound was not observed in a parallel experiment in which AN was replaced with sodium nitrate (SN) (Fig. S6a; see Sect. 3.3 for discussion). The potential aqSOA formation ~~most probable~~ pathways via ~~of direct photosensitized and nitrate-mediated photo-oxidation of~~ VL photo-oxidation in this study are summarized in Fig. 2~~were proposed (Fig. 4). In Scheme 1 (pH 4 and pH <4 under air-saturated conditions), ^{3}VL* and $^{\bullet}$OH (from ^{3}VL* or nitrate photolysis) can initiate H-atom abstraction to generate phenoxy or ketyl radicals (Huang et al., 2018; Vione et al., 2019).~~ At pH 4, ring-opening products (Fig. S5) from fragmentation in both VL* and VL+AN may have reacted with VL or dissolved ammonia to generate $C_{10}H_{10}O_5$ (No. 4, Table S2) (Pang et al., 2019b) or~~and~~ a potential imidazole derivative ($C_5H_5N_3O_2$; No. 5, Table S2), respectively. Moreover, nitrate photolysis products promoted functionalization and nitration (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S2). ~~At pH<4, the reactivity of ^{3}VL* increased as suggested by the abundance of oligomers (e.g., $C_{16}H_{14}O_6$) and increased normalized abundance of N-containing compounds.~~

Line 364: At pH < 4, ^{3}VL* likely have higher reactivity as suggested by the increased normalized abundance of oligomers (e.g., $C_{16}H_{14}O_6$; No. 6, Table S2 and $C_{31}H_{24}O_{11}$) and N-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S2 and $C_{13}H_{14}N_2O_{10}$) (Fig. 2).

Minor Comments

1. Line 25. This notion of "efficiency" (i.e., which reaction path is faster) depends on the concentrations of the two oxidant precursors, VL and nitrate. Thus it's not a universally true statement.

Response: Thank you for this suggestion. The apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of either VL or nitrate during simulated sunlight illumination can be defined as (Anastasio et al., 1996; Smith et al., 2014, 2016):

$$\Phi_{\mathrm{GUA}} = \frac{\text{mol GUA destroyed}}{\text{mol photons absorbed}} \qquad \text{(Eq. S9)}$$

Φ_{GUA} was calculated using the measured rate of GUA decay and rate of light absorption by either VL or nitrate through the following equation:

$$\Phi_{\mathrm{GUA}} = \frac{\text{rate of GUA decay}}{\text{rate of light absorption by VL or nitrate}} = \frac{k'_{\mathrm{GUA}} \times [\mathrm{GUA}]}{\sum[\left(1-10^{-\varepsilon_\lambda [\mathrm{C}] l}\right) \times I'_\lambda)} \qquad \text{(Eq. S10)}$$

where k'_{GUA} is the pseudo-first-order rate constant for GUA decay, [GUA] is the concentration of GUA (M), ε_λ is the base-10 molar absorptivity (M^{-1} cm^{-1}) of VL or nitrate at wavelength λ, [C] is the concentration of VL or nitrate (M), l is the pathlength of the illumination cell (cm), and I'_λ is the volume-averaged photon flux (mol-photons L^{-1} s^{-1} nm^{-1}) determined from 2NB actinometry:

$$j(\mathrm{2NB}) = 2.303 \times \Phi_{\mathrm{2NB}} \times l \times \Sigma_{300\,nm}^{350\,nm} (\varepsilon_{\mathrm{2NB},\lambda} \times I'_\lambda \times \Delta\lambda) \qquad \text{(Eq. S11)}$$

The ϕ_{GUA} in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL. We have revised this in the text as follows and added the information shown above in the supporting information: **Text S7**. Estimation of the apparent quantum efficiency of guaiacol photodegradation.

Line 29: Furthermore, comparisons of the apparent quantum efficiency of guaiacol photodegradation indicate that in this study, guaiacol oxidation by photosensitized reactions of VL ~~was observed to be more~~ is less efficient relative to nitrate-mediated photo-oxidation.
Other relevant revisions in the text are as follows:

Line 514: As mentioned earlier, ~~the~~ ^{3}VL* chemistry appears to be more important than that of nitrate photolysis even at 1:10 molar ratio of VL/nitrate on account of the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. However, the apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL (see Text S7 for the more details).

Line 619: The oxidation of guaiacol, a non-carbonyl phenol, by photosensitized reactions of vanillin was also shown to be ~~more~~ less efficient than that by nitrate photolysis products based on its lower apparent quantum efficiency.

2. L. 42. "respectively" doesn't serve a purpose in this sentence.

Response: Agree, we have deleted 'respectively' from line 50 (formerly 42).

3. Section 2.1. What was the initial volume of solution illuminated? Were solutions stirred? What was the flow rate of gas (N2 or air) through the solution before and during the experiment?

Response: The initial volume of the illuminated solution is 500 mL. The solutions were continuously mixed throughout the experiments. A constant flow rate of 0.5 dm^3/min was used before and during the experiments. These have been added to the text as follows:

Line 102: Photo-oxidation experiments were performed in a ~~500-mL~~ custom-built quartz photoreactor. The solutions (initial volume of 500 mL) were continuously mixed throughout the experiments using ~~equipped with~~ a magnetic stirrer. The solutions were bubbled with synthetic air or nitrogen (N_2) (> 99.995%) (0.5 dm^3/min) for 30 min before irradiation to achieve air- and N_2-saturated conditions, respectively, and the bubbling was continued throughout the reactions (Du et al., 2011; Chen et al., 2020).

4. L. 100. Was there a difference in the temperature between the illuminated and dark solutions?

Response: For all experiments, the range in temperature fluctuations was 27 ± 2 °C.

5. L. 106. If the authors are going to abbreviate 2-propanol as IPA, it would be better to call it isopropyl alcohol to help readers remember the name of the abbreviation. NaBC is a poor choice for an abbreviation for sodium bicarbonate since BC stands for black carbon typically. Better to simply use its chemical formula, NaHCO3 or HCO3-, depending on the context.

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

6. L. 111. 2-propanol and bicarbonate were added in some experiments, but the description of why is odd. Their primary role will be OH scavengers, so it's strange to call them a VOC and

inorganic anion, respectively. 2-propanol is not a common atmospheric gas, so it's a poor choice of model VOC. Similarly, calling bicarbonate an "inorganic anion" is a poor choice of words, since sulfate and nitrate are the classic inorganic ions. Better to refer to 2-propanol and bicarbonate as "OH scavengers" since that is their main role.

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

7. L. 113. What does it mean that the OH scavengers were not added "in excess"? Since they're reacting with OH (which will have a very low concentration) they are technically in excess. Better to avoid this discussion, as it's not fruitful. If you want to dive more into the OH scavengers, you could calculate the fraction of OH each intercepts in their respective solutions or the amount that they suppress the OH concentration. (But, again, this depends on if the species were purged from solution.)

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

8. L. 151. The disproportionation of HO_2/O_2^- is the same as the reaction of HO_2 with O_2^-, so this sentence repeats itself.

Response: The reviewer is right. We have corrected this sentence as follows:

Line 181: The disproportionation of $HO_2^{\bullet}/O_2^{\bullet-}$ (Anastasio et al., 1996) ~~and reaction of $HO_2^{\bullet}$ with $O_2^{\bullet-}$ (Du et al., 2011)~~ form hydrogen peroxide (H_2O_2), which is a photolytic source of $^{\bullet}OH$.

9. L. 163. It's unclear what the authors mean by "...a minimal role for $^3VL^*$ in VL photo-oxidation". Do they mean that $^3VL^*$ + VL is an unimportant reaction (but see above about this) or that the direct photodegradation of VL doesn't proceed through the triplet state?

Response: In this study, the photodegradation of VL is mainly governed by VL triplets, as explained in our response to major comment #2 (Please see this response for the revised text).

10. L. 167. It is not true that $^1O_2$ has a much longer lifetime than $^3C^*$; rather, the lifetimes are approximately the same. In cloud and fog drops, the lifetime of $^1O_2$ is controlled by water deactivation and is approximately 5 us (see Bilski et al., 1997). The lifetime of $^3C^*$is controlled by reaction with dissolved O2 and is approximately 1/((2E9 M-1 s-1)*(250uM)) ~ 2 us. Also, rather than the oxidant lifetime, it is the product of the oxidant concentration times its second-order rate constant that determines the relative importance of a given oxidant.

Response: We thank the reviewer for the correction. This sentence has been modified as part of the revision for major comment #2 (Please see this response for the revised text).

11. Page 7. This whole page is one paragraph. It should be trimmed to reduce speculative discussions of mechanisms, then broken into smaller pieces, focused on certain themes/points.

Response: We agree with the reviewer. Section 3.1.1 has been revised to reduce speculative discussions of mechanisms (Revisions for lines 188–220 are shown in response to major comment #2). Changes in the text are as follows:

Lines 220–240: ~~Similar to VL*, t~~The decay rate constant for VL+AN under air-saturated conditions was also higher ~~faster~~ (6.6 times) than N_2-saturated conditions~~,~~ which may c~~an~~ be due to ~~several~~ reactions facilitated by nitrate photolysis products ~~and the enhancement of N(III) mediated photo-oxidation~~ that may have been enhanced in the presence of O_2 ~~as reported in early works~~ (Vione et al., 2005; Kim et al., 2014; Pang et al., 2019a). As shown later, more nitrogen-containing species were observed under air-saturated conditions. An example is enhanced VL nitration likely from increased $^{\bullet}NO_2$ formation such as from the reaction of $^{\bullet}OH$ and $O_2^{\bullet-}$ with NO_2^- (Reactions 4 and 5, respectively; Table 1) or the autoxidation of $^{\bullet}NO$ from NO_2^- photolysis (Reactions 6–9; Table 1) in aqueous solutions (Pang et al., 2019a). ~~Reactions involving $^{\bullet}HO_2/O_2^{\bullet-}$ which may originate from the photolysis of nitrate alone, likely from the production and subsequent photolysis of peroxynitrous acid (HOONO) (Reaction 10; Table 1) (Jung et al., 2017; Wang et al., 2021), or the reactions of ^{3}VL* in the presence of O_2, may have contributed as well. For instance, Wang et al. (2021) recently demonstrated that nitrate photolysis generates $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and $HONO_{(g)}$ in the presence of dissolved aliphatic organic matter (e.g., nonanoic acid, ethanol), with the enhanced $HONO_{(g)}$ production caused by secondary photochemistry between $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and photoproduced $NO_{x(aq)}$ (Reactions 11 and 12; Table 1), in agreement with Scharko et al. (2014). The significance of this increased HONO production is enhanced $^{\bullet}OH$ formation (Reaction 13; Table 1). In addition, $^{\bullet}HO_2$ can react with $^{\bullet}NO$ (Reaction 10; Table 1) from NO_2^- photolysis (Reaction 6; Table 1) to form HOONO, and eventually $^{\bullet}NO_2$ and $^{\bullet}OH$ (Reaction 14; Table 1) (Pang et al., 2019a)~~. Nevertheless, the comparable decay rate~~s~~ constants for VL* and VL+AN imply that ^{3}VL* chemistry still dominates even at 1:10 molar ratio of VL/nitrate, probably due to the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. Although we have no quantification of the oxidants in our reaction systems as it is outside the scope of this study, these observations clearly substantiate that~~secondary oxidants from ^{3}VL*, which are formed when O_2 is present, are required for efficient~~ photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation are more efficient in the presence of O_2.

Lines 241–300: The products from VL* under N_2-saturated conditions were mainly oligomers (e.g., $C_{16}H_{14}O_4$) (Fig. 1a), consistent with triplet-mediated oxidation forming higher molecular weight products, probably with less fragmentation relative to oxidation by $^{\bullet}OH$ (Yu et al., 2014; Chen et al., 2020). A threefold increase in the normalized abundance of products was noted upon addition of nitrate (VL+AN under N_2-saturated conditions; Fig. 1b), and in addition to oligomers, functionalized monomers (e.g., $C_8H_6O_5$) and nitrogen-containing compounds (e.g., $C_8H_9NO_3$; No. 2, Table S2~~3~~) were also observed, in agreement with $^{\bullet}OH$-initiated oxidation yielding more functionalized/oxygenated products compared to triplet-mediated oxidation (Yu et al., 2014; Chen et al., 2020). ~~Compared to N_2-saturated conditions, the normalized abundance of products such as o~~Oligomers, ~~and~~ functionalized monomers (e.g., demethylated VL; Fig. S4), and nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3~~3,~~ ~~Table S3~~,Table S2) (for VL+AN) had higher normalized abundance ~~were also more relatively abundant~~ under air-saturated conditions ~~were~~

~~significantly higher under air-saturated conditions~~ (Figs. 1c-d), likely due to efficient ~~the secondary oxidants from~~ ^{3}VL*-initiated oxidation and enhanced VL nitration in the presence of ~~and~~ O_2 ~~and their interactions with nitrate photolysis products~~. ~~The nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S3) were also more relatively abundant under air-saturated conditions.~~ For both VL* and VL+AN under air-saturated conditions, the most abundant product was $C_{10}H_{10}O_5$ (No. 4, Table S2~~3~~), a substituted VL. Irradiation of VL by 254-nm lamp has also been reported to lead to VL dimerization and functionalization via ring-retaining pathways, as well as small oxygenates but only when $^{\bullet}$OH from H_2O_2 were involved (Li et al., 2014). In this work, small organic acids (e.g., formic acid) were observed from both VL* and VL+AN under air-saturated conditions (Fig. S5) due to simulated sunlight that could access the 308-nm VL band (Smith et al., 2016). Interestingly, we observed a potential imidazole derivative ($C_5H_5N_3O_2$; ~~Fig. 1d~~No. 5, Table S2) from VL+AN under air-saturated conditions (Fig. 1d), which may have formed from reactions induced by ammonium. This compound was not observed in a parallel experiment in which AN was replaced with sodium nitrate (SN) (Fig. S6a; see Sect. 3.3 for discussion). The potential aqSOA formation ~~most probable~~ pathways via ~~of direct photosensitized and nitrate-mediated photo-oxidation of~~ VL photo-oxidation in this study are summarized in Fig. 2~~were proposed (Fig. 4). In Scheme 1 (pH 4 and pH <4 under air-saturated conditions), ^{3}VL* and $^{\bullet}$OH (from ^{3}VL* or nitrate photolysis) can initiate H-atom abstraction to generate phenoxy or ketyl radicals (Huang et al., 2018; Vione et al., 2019).~~ At pH 4, ring-opening products (Fig. S5) from fragmentation in both VL* and VL+AN may have reacted with VL or dissolved ammonia to generate $C_{10}H_{10}O_5$ (No. 4, Table S2) (Pang et al., 2019b) or~~and~~ a potential imidazole derivative ($C_5H_5N_3O_2$; No. 5, Table S2), respectively. Moreover, nitrate photolysis products promoted functionalization and nitration (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S2). ~~At pH<4, the reactivity of ^{3}VL* increased as suggested by the abundance of oligomers (e.g., $C_{16}H_{14}O_6$) and increased normalized abundance of N-containing compounds.~~

The molecular transformation of VL upon photo-oxidation was examined using the van Krevelen diagrams (Fig. S7). For all experiments (A1-15~~9~~; Table ~~S~~2) in this study, the O:C and H:C ratios of the products were mainly similar to those observed from the aging of other phenolics (Yu et al., 2014) and BB aerosols (Qi et al., 2019). U~~u~~nder N_2-saturated conditions, o~~O~~ligomers with O:C ratios ≤ 0.6 were dominant in VL*, ~~under N_2-saturated conditions.~~while smaller molecules (n_c ≤ 8) with higher O:C ratios (up to 0.8) were also observed f~~F~~or VL+AN. ~~For VL+AN under air-saturated conditions, smaller molecules (n_c ≤8) with higher O:C ratios (up to 0.8) were also observed.~~ In contrast, ~~M~~more products with higher O:C ratios (≥ 0.6) were noted under air-saturated conditions for both VL* and VL+AN. F~~f~~or experiments A5 to A8, ~~The~~ H:C ratios were mostly around 1.0 and double bond equivalent (DBE) values were typically (58% of the 50 most abundant products) > 7, indicating that the products ~~for experiments A5 to A8 (Table S2)~~ were mainly oxidized aromatic ~~aromatic species~~ compounds (Xie et al., 2020). Compounds with H:C ≤ 1.0 and O:C ≤ 0.5 are common for aromatic species, while compounds with H:C ≥ 1.5 and O:C ≤ 0.5 are typical for more aliphatic species (Mazzoleni et al., 2012; Kourtchev et al., 2014; Jiang et al., 2021). ~~Moreover, majority of the products for experiments A5 to A8 have double bond equivalent (DBE) values >7, which corresponds to oxidized aromatic compounds (Xie et al., 2020).~~ In contrast, Loisel et al. (2021) reported mainly oxygenated aliphatic-like compounds ~~(H:C, ≥ 1.5 and O:C, ≤ 0.5 ratios)~~ from the direct irradiation of VL (0.1 mM), ~~which may be~~ probably due to

their use of ESI in the negative ion mode, which has higher sensitivity for detecting compounds such as carboxylic acids (Holčapek et al., 2010; Liigand et al., 2017), and solid-phase extraction (SPE) procedure causing the loss of some oligomers (LeClair et al., 2012; Zhao et al., 2013; Bianco et al., 2018). Among experiments A5 to A8 ~~(Table S2)~~, VL+AN under air-saturated conditions (A7) had the highest normalized abundance of products and <OS_c>, ~~most~~ probably due to the combined influence of ~~the secondary oxidants from~~ ^{3}VL* and enhanced VL nitration in the presence of O_2~~, and nitrate photolysis products~~. ~~In our calculations, the increase in <OS_c> (except for VOCs and inorganic anions experiments; A9 to A12; Table S2) was lower than those in $^{\bullet}$OH- or triplet-mediated oxidation of phenolics (e.g., phenol, guaiacol) measured using an aerosol mass spectrometer (Sun et al., 2010; Yu et al., 2014),~~ Our measured <OS_c> range from -0.28 to +0.12, while other studies on phenolic aqSOA formation reported <OS_c> ranging from -0.14 to +0.47 (Sun et al., 2010) and 0.04 to 0.74 (Yu et al., 2014). This is likely because we excluded contributions from ring-opening products, which may have higher OS_c values as these products are not detectable in the positive ion mode. Thus, the <OS_c> in this study likely were lower estimates. In brief, the presence of ~~secondary oxidants from ^{3}VL* and~~ O_2 increased the normalized abundance of products and promoted the formation of more oxidized aqSOA. ~~These trends were reinforced in the presence of nitrate, indicating synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ Compared to N_2-saturated condition, the higher normalized abundance of nitrogen-containing products under air-saturated condition for VL+AN (at pH 4) suggests a potential enhancement of VL nitration in the presence of O_2.

12. L. 196. The text here and elsewhere discusses the abundance of specific products (not just the normalized product abundance). The abundance of each product should be added to Table S3, along with some estimate of the relative uncertainty of these values.

Response: The abundance mentioned in line 249 (formerly 196) pertains to the normalized abundance of all nitrogen-containing products for VL* at pH 4 under air-saturated conditions, not for specific products.

13. L. 224. How much lower are OS(C) values here compared to previous work on aqSOA? Compare these values.

Response: Our measured <OS_c> range from -0.28 to +0.12, while other studies on phenolic aqSOA formation reported <OS_c> ranging from -0.14 to +0.47 (Sun et al., 2010) and 0.04 to 0.74 (Yu et al., 2014). This information has been added to the text as follows:

Line 286: Our measured <OS_c> range from -0.28 to +0.12, while other studies on phenolic aqSOA formation reported <OS_c> ranging from -0.14 to +0.47 (Sun et al., 2010) and 0.04 to 0.74 (Yu et al., 2014). This is likely because we excluded contributions from ring-opening products, which may have higher OS_c values as these products are not detectable in the positive ion mode.

14. L. 240. Is there any evidence that ^{3}VL* + O_2 directly makes OH? This would seem energetically unfavorable and also to be minor compared to energy transfer to make $^1O_2$.

Response: We do not have direct evidence of $^{\bullet}OH$ formation from $^{3}VL^{*} + O_2$, although trace amounts of H_2O_2 were likely formed during VL photodegradation (Li et al., 2014) similar to the case of other phenolic compounds (Anastasio et al., 1996). This statement has been added to line 241.

Line 316: Trace amounts of H_2O_2 were likely formed during VL photodegradation (Li et al., 2014), similar to the case of other phenolic compounds (Anastasio et al., 1996).

15. L. 242. In the presence of O_2, the ketyl radical is probably too short lived (it reacts with O2 to make an alpha-hydroxy peroxyl radical) to combine appreciably with a phenoxyl radical. But the phenoxyl radical is in resonance with a carbon-centered cyclohexadienyl radical that is longer lived; these two species can couple (Yu et al., ACP, 2014).

Response: We thank the reviewer for the additional information. We have now revised line 311 to include this as follows:

Line 317: Oligomers can then form via the coupling of phenoxy radicals or phenoxy and cyclohexadienyl ~~ketyl~~ radicals (Sun et al., 2010; Yu et al., 2014~~Berto et al., 2016~~; Vione et al., 2019).

16. L. 261. The rate constant for H_2O_2 formation is fastest near the pKa of HO_2, i.e., pH 4.8, so one wouldn't expect greater H_2O_2 formation at pH 2.5 compared to pH 4. But this also depends on the pH dependence of the HO_2/O_2^- sources and sinks.

Response: Thank you for pointing this out. This statement has been deleted.

17. L. 264. This discussion of the pH dependence of N(III) photolysis doesn't seem applicable since the addition of nitrate makes a negligible contribution to VL decay. Just because N(III) photolysis is pH dependent doesn't mean it matters here.

Response: Thank you for pointing out this error. The reviewer is correct that the pH dependence of N(III) is not relevant for the discussion of pH effects on VL decay rate constants. Some of these statements have been transferred to the next paragraph (from line 360) to provide a potential explanation for the increased normalized abundance of nitrogen-containing compounds at lower pH. Revisions for lines 333–354 are shown in our response to major comment #6. Other changes in the text are as follows:

Line 359–364: For VL+AN, the normalized abundance of nitrogen-containing compounds was also higher ~~increased~~ at lower pH (Table ~~S~~2), likely due to increased $^{\bullet}OH$ and $^{\bullet}NO_2$ formation, which may be caused by the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a). At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}OH$ formation than that of NO_2^- in the near-UV region

(Arakaki et al., 1999; Kim et al., 2014). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}OH$ (Reactions 4 and 10~~5~~; Table 1) (Pang et al., 2019a).

Line 372: Essentially, the higher reactivity of $^3VL^*$ and predominance of HONO over nitrite at lower pH may have resulted in increased formation of products mainly composed of oligomers and functionalized monomers.

18. L. 299. Why would the presence of HO2 lead to more dimer formation? HO2 (and O2-) are too weak to oxidize phenols at any significant rate.

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

19. Lines 298-301: This argument is circular: IPA cannot make more OH by scavenging OH and turning it into HOOH, which then photolyzes to make OH. Think of the associated stoichiometry. IPA will suppress [OH] because it is an OH sink, thus rendering OH an insignificant oxidant for VL. The observation that IPA has a negligible impact on VL decay (Fig. S3c) indicates that OH is not important as an oxidant for VL (with or without IPA) or that the IPA was mostly purged from the system.

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

20. L. 310. IPA makes no difference in the VL kinetics, whether nitrate is present or not. So please don't make sweeping statements such as "...the role of nitrate in VL photo-oxidation is enhanced in the presence of IPA...". And don't suggest that OH is an important intermediate in the formation of a product in the presence of IPA (e.g., line 311), since IPA will greatly suppress the OH concentration.

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

21. L. 313-322. It is hard to believe that 1 mM of IPA can significantly disrupt the structure of 55 M water. In any case, there is no increase in light absorption by the aqSOA formed in the presence of IPA (Fig. 2), so the Berke mechanism seems unimportant. Most of this should be deleted.

Response: Sentences related to $^{\bullet}OH$ scavengers have been deleted from the original manuscript.

22. L. 327. It is difficult to imagine that carbonate radical is a significant oxidant in these experiments: carbonate rate constants are relatively slow (compared to triplets or OH) and VL photodegradation is very fast. If the authors want to propose carbonate radical as an important sink, they need to do some calculations of its steady-state concentration and estimate the corresponding rate of VL loss. Again, the qualitative normalized abundance of products is driving these uncertain statements, while the quantitative photodecay rates and light absorption are showing there is no significant effect of bicarbonate. Lead with the latter observations, as they are more robust.

Response: Sentences related to •OH scavengers have been deleted from the original manuscript.

23. L. 336. 1 mM IPA or bicarbonate is not high enough to reduce the cage effect from nitrate photolysis. In any case, IPA or bicarbonate are OH sinks, so they will suppress, not enhance, the OH concentration.

Response: Sentences related to •OH scavengers have been deleted from the original manuscript.

24. L. 376. If this proposed mechanism was true, then VL decay would be significantly faster in the presence of nitrate, but this is not the case. It's not clear what the authors are trying to explain here - is it the increase in oligomerization at higher [VL]? The explanation for this is probably that the concentrations of phenoxyl radicals (and the related, carbon-centered cyclohexadienyl radicals) increase with [VL], making radical-radical recombination to form oligomers a more significant fate.

Response: Thank you for pointing this out. We apologize for the confusion. There should have been an explanation here for the increased oligomerization observed at higher [VL]. We agree with the reviewer that a possible reason for this is the increased concentration of radicals with [VL] (added to the revised text). The succeeding lines are for comparing the potential pathways at 1:1 VL/nitrate and 1:100 VL/nitrate only at low [VL]. The text has been revised to clarify these as follows:

Lines 466–485: For both VL* and VL+AN, the contribution of < 200 m/z to the normalized abundance of products was higher at low [VL] than at high [VL], while the opposite was observed for > 300 m/z (Fig. S12~~3~~d). This indicates that functionalization was favored at low [VL], as supported by the higher $<OS_c>$, while oligomerization was the dominant pathway at high [VL], consistent with more oligomers or polymeric products reported from high phenols concentration (e.g., 0.1 to 3 mM) (Li et al., 2014; Slikboer et al., 2015; Ye et al., 2019). This is probably due to an increased concentration of phenoxy radicals (in resonance with a carbon-centered cyclohexadienyl radical) at high [VL], promoting radical-radical polymerization (Sun et al., 2010; Li et al., 2014). ~~Moreover~~ At low [VL], the contribution of < 200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, ~~further~~ suggesting the prevalence of functionalization for the former ~~formation of more oxidized products~~. ~~This may also be the reason why~~ In addition, 1:1 VL/nitrate (A11~~5~~; Table ~~S~~2) had higher $<OS_c>$ than 1:100 VL/nitrate (A12~~6~~; Table ~~S~~2), indicating the formation of more oxidized products, but had fewer N-containing compounds compared to the latter. A possible explanation is that at 1:1 VL/nitrate, VL may ~~efficiently~~ compete~~s~~ with NO_2^- for •OH (from nitrate or nitrite photolysis, Reaction 4; Table 1) and indirectly reduce~~s~~ $^{\bullet}NO_2$. Similarly, hydroxylation has been suggested to be a~~n~~ more important pathway for 1:1 VL/nitrite than in 1:10 VL/nitrite (Pang et al., 2019a). ~~This may also be the reason why 1:1 VL/nitrate (A15; Table S2) had higher $<OS_c>$ than 1:100 (A16; Table S2) VL/nitrate but had fewer N-containing compounds compared to the latter.Moreover, the contribution of <200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, further suggesting the formation of more oxidized products.~~ Fragmentation, which leads to the decomposition of previously formed oligomers and generation

of small, oxygenated products such as organic acids (Huang et al., 2018), may also occur for the low [VL] experiments. However, its importance would likely be observed at longer irradiation times, similar to the high [VL] experiments.

25. L. 405. GUA should not undergo any direct photochemistry, so its decay in the absence of VL or AN suggests either that there is an oxidant-making contaminant in the system (that is consumed within a few hours) or that GUA is evaporating during illumination. But there is no GUA loss in the dark: is this because the temperature was cooler in the dark?

Response: Vione et al. (2019) also observed the direct photodegradation of GUA (0.1 mM) upon irradiation using Xe lamp. Moreover, Sun et al. (2010) reported an intermediate rate of direct photoreaction for GUA (0.1 mM), yielding aqSOA including GUA dimers (similar to what we observed). There was no significant loss of GUA both after 30 min of purging in the dark (-0.36%) and after 360 min of dark control experiments (-1.7%). Also, the temperature fluctuations (27 ± 2 °C) were minimal.

26. L. 498. This sentence mentions "Further enhancement of VL photo-oxidation…in the presence of nitrate…", but VL photo-oxidation (i.e., photodegradation) was not enhanced in the presence of nitrate.

Response: This statement refers to the presence of both O_2 and nitrate resulting in the highest normalized abundance of products (including N-containing compounds) and $<OS_c>$ among experiments A5–A8. We have revised this as follows:

Line 606–611: ~~Further enhancement of VL photo-oxidation under air-saturated conditions in the presence of nitrate indicates synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ In contrast, ^{3}VL*-initiated reactions proceeded rapidly under air-saturated conditions (O_2 is present) as indicated by higher VL decay rate constant and increased normalized abundance of products. For pH 4 experiments, the presence of both O_2 and nitrate resulted in the highest normalized abundance of products (including N-containing compounds) and $<OS_c>$, which may be due to the presence of O_2 promoting VL nitration.

27. L. 1004. The author order is incorrect on the Tinel et al. ref.

Response: Thank you for the correction. We have now amended the author order for this reference as follows and revised the corresponding in-text citations:

George, C., Brüggemann, M., Hayeck, N., Tinel, L., and Donaldson, D. J.: Interfacial photochemistry: physical chemistry of gas-liquid interfaces, in: Developments in Physical & Theoretical Chemistry, edited by: Faust, J. A. and House, J. E., Elsevier, 435–457, https://doi.org/10.1016/B978-0-12-813641-6.00014-5, 2018.

28. Table 1. The quantum yield for Rxn 3 is not 0.001. This is a misperception based on the O(3P) result of Warneck and Wurzinger (J Phys Chem, 1988); their paper shows a value of ~0.01 for more direct (nitrite) measurements. Benedict et al. (Env Sci Technol., 2017) confirmed this higher

value. This error doesn't affect the current work, but it would be a shame to propagate the misperception.

Response: We thank the reviewer for the careful read. We now have corrected this value in Table 1:

$NO_3^- + hv \rightarrow NO_2^- + O(^3P)$; $\varphi = 0.\cancel{0}011$

29. Figure 4. (a) The resolution of the figure is poor, so it's fuzzy and hard to read. (b) Scheme 1 suggests that oligomers are only formed at pH < 4, which isn't true, as past work has shown oligomer formation in similar phenol systems at pH 5. (c) Ketyl radicals formed by $^3C^*$ + phenol typically are shown as phenoxyl OH group (a result of the triplet abstracting a hydrogen) and no double bond between the C and O. As stated earlier, their lifetimes are short in the presence of O2, so they're unlikely to do the coupling as shown here.

Response:

(a) This may be a formatting issue. We ensured that this is avoided in the revised version.

(b) The reviewer is right that oligomer formation in similar phenol systems was also observed at pH 5. However, the molecular formulas initially presented in Fig. 4 were the most abundant products or products with a significant increase in normalized abundance. To avoid this confusion, Fig. 4, now 2, has been revised to show the major products for each condition, with a marker for the most abundant products.

(c) Thank you for catching this error. Ketyl radical has been deleted in Fig. 2.

Supplemental Material Notes

1. General note – it would have been helpful to have line numbers in the supplement.

Response: Line numbers have now been added to the supporting information as well.

2. Text S3. Were calibration curves only made once? Were they actually used in quantifying VL and GUA? (I don't see the need since absolute values are not needed in the kinetic plots.)

Response: No, the calibration curves were prepared weekly to account for potential changes in the detector response of UHPLC. These calibration curves were used to quantify VL and GUA as the calculated VL concentration was used for estimating the normalized abundance of products.

3. Text S6. (a) It's unclear what is meant by "Then, the average relative intensity absorbed by 2NB solution as a function of wavelength was calculated." Can you show this with an equation? (b) How much did the photon flux vary between experiments? Was this determined? If not, this variation is a source of variability in the kinetic measurements.

Response: (a) We apologize for the confusion. This statement pertains to a scaling factor (SF) that was used to determine the absolute photon flux in the reactor, I'_λ. Similar to Smith et al. (2014, 2016), we measured the spectral shape of the photon output of our illumination system (i.e., the relative flux at each wavelength) using a high-sensitivity spectrophotometer (Brolight Technology Co. Ltd, Hangzhou, China). Using an SF, this measured relative photon output, $I_\lambda^{\text{relative}}$, is related to I'_λ as follows:

$$I'_\lambda = I_\lambda^{\text{relative}} \times \text{SF} \qquad \text{(Eq. S6)}$$

Substitution of Eq. S6 into Eq. S5 and rearrangement yields:

$$j(\text{2NB}) = 2.303 \times (10^3\ \text{cm}^3\ \text{L}^{-1} \times\ 1\ \text{mol/N}_\text{A}\ \text{mlc}) \times \sum\left(I'_\lambda \times \Delta\lambda \times \varepsilon_{\text{2NB},\lambda} \times \Phi_{\text{2NB}}\right) \qquad \text{(Eq. S5)}$$

Where j(2NB) is the 2NB decay rate constant, N_A is Avogadro's number, I'_λ is the actinic flux (photons cm^{-2} s^{-1} nm^{-1}), $\Delta\lambda$ is the wavelength interval between actinic flux data points (nm), and $\varepsilon_{\text{2NB},\lambda}$ and $\Phi_{\text{2NB},\lambda}$ are the base-10 molar absorptivity (M^{-1} cm^{-1}) and quantum yield (molecule $photon^{-1}$) for 2NB, respectively. Values of $\varepsilon_{2NB,\lambda}$ (in water) at each wavelength under 298 K and a wavelength-independent Φ_{2NB} value of 0.41 were adapted from Galbavy et al. (2010).

$$\text{SF} = \frac{j(\text{2NB})}{2.303 \times (10^3\ \text{cm}^3\ \text{L}^{-1} \times 1\ \text{mol/N}_\text{A}\ \text{mlc}) \times \sum\left(I_\lambda^{\text{relative}} \times \Delta\lambda \times \varepsilon_{\text{2NB},\lambda} \times \Phi_{\text{2NB}}\right)} \qquad \text{(Eq. S7)}$$

and substitution of (Eq. S6) into (Eq. S7) yields:

$$I'_\lambda = I_\lambda^{\text{relative}} \frac{j(\text{2NB})}{2.303 \times (10^3\ \text{cm}^3\ \text{L}^{-1} \times 1\ \text{mol/N}_\text{A}\ \text{mlc}) \times \sum\left(I_\lambda^{\text{relative}} \times \Delta\lambda \times \varepsilon_{\text{2NB},\lambda} \times \Phi_{\text{2NB}}\right)} \qquad \text{(Eq. S8)}$$

Finally, $I'_\lambda$was estimated through Eq. S8~~5~~. The estimated photon flux in the aqueous reactor is shown in Figure S1~~2~~.

We have added the above information to Text S6.

(b) The j(2NB) in this study varied from 0.0021 to 0.0026 s^{-1}. The decay rate constants have now been normalized to the photon flux, and the updated values are shown in Table 2, formerly S2.

4. Table S2. (a) VL (and GUA) decays are rate constants, not decay rates. (b) For reference, it would be helpful to give the OS(C) of VL. (c) What is pH of expt. A19?

Response: (a) Thank you for catching this error. This has been corrected in Table S2, now 2, as well as elsewhere in the text:

Table ~~S~~2. Reaction conditions, initial VL (and GUA) decay rate~~s~~ constants

Line 142: the calculation of GUA decay rate constant.

Line 172: initial VL (and GUA) decay rate~~s~~ constants,

Line 204: Contrastingly, the VL* decay rate constant under air-saturated conditions

Line 220: ~~Similar to VL*~~, ~~t~~The decay rate constant for VL+AN

Line 235: Nevertheless, the comparable decay rate~~s~~ constants

Line 331: The decay rate~~s~~ constants

Line 509: GUA decay rate constant was ~~faster~~ higher by 2.2 (GUA+VL)

Line 512: This enhanced GUA decay rate constant

Line 574: Although nitrate did not substantially affect the VL decay rate~~s~~ constants,

(b) Agree, we have now added the OS_c of VL (-0.25) to Table 2.

(c) The pH for exp 19, now 15, is 4, same with other experiments involving GUA. This is already listed in the second column of Table 2.

5. Figure S1. The vanillin spectrum has a problem around 305 nm - a large discontinuity that is probably caused by lamp switch. Either reacquire the spectrum or replace with a published value.

Response: Thank you for catching this. Fig. S1 has been revised to correct this.

6. Figure S3. Were the decays ever determined multiple times for the same condition? It would be helpful to show these results and derive a relative uncertainty for decay rate constants.

Response: Yes, the decays reported in Fig. S3 are the average of results from triplicate experiments, and the error bars for each data point are already shown. The photon flux-normalized decay rate constants have also been updated in Table S2, now 2, along with the standard deviation for each condition (Please see our response to major comment #5). The following sentence has been added in the methods section to clarify this:

Line 132: Each experiment was repeated independently at least three times and measurements were done in triplicate. The reported decay rate constants and absorbance enhancement are the average of results from triplicate experiments, and the corresponding errors represent one standard deviation.

We also added this note to figure captions when applicable: most error bars are smaller than the markers.

7. Figure S6. How can we tell that the imidazole formed in the AN experiment was not formed in the SN experiment? It would be helpful to put a marker on the two plots of Fig. S6 to show where the imidazole showed up in the AN experiment.

Response: Thank you for pointing this out. Fig. S6 has been revised to show a marker for the potential imidazole compound.

Recommendation

I recommend that the manuscript be majorly revised and then reconsidered.

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Author Response for "Aqueous SOA formation from the photo-oxidation of vanillin: Direct photosensitized reactions and nitrate-mediated reactions" by Mabato et al.

We thank the reviewer for the thorough review and many constructive comments that helped improve the manuscript. Our point-by-point responses are below (changes to the original manuscript text and supporting information are in red, moved content in double-line strikethrough, and removed content in strikethrough). Please note that the line numbers in the responses refer to our revised manuscript with tracked changes. Also, please note that because we restructured the manuscript, the numbering of some figures and tables in the revised manuscript is different from those in the original manuscript.

Reviewer 2

The manuscript describes very well-designed studies of vanillin photooxidation in bulk liquid solutions where pH, concentrations, reactant ratios, dissolved gases (N2 or O2), ions (nitrate, bicarbonate) and other species (isopropanol) were varied in many combinations. The work is technically sound, with the loss of reactants, the identification and quantification of products, and the absorbance changes in solution all monitored hourly. The authors exhaustively discuss the differences between each experimental variation, pulling out as much detail as possible. This paper will be of interest to those interested in biomass burning aerosol and brown carbon formation, and is publishable after major revision to address the following points.

1. In places the discussion veers off into speculation, or suggests theories that aren't adequately explained enough to be convincing to the reader, as noted below. Generally the discussion is convincing and well-connected to the literature, but the discussion section reads like it has a thousand detailed conclusions, leaving the reader often feeling "lost in the weeds" and blunting the impact of the work. In general, the focus of the paper could be improved by moving Table 1 to the SI, removing a lot of speculative discussion, and bringing Tables S2 and maybe S3 from the SI to the main paper. These tables are more vital to the discussion at many points, in my opinion.

Response: Thank you for the suggestion. Speculative discussions have been removed from the revised text. First, Section 3.1.1 has been amended to emphasize the importance of VL triplets:

Line 187: ~~Effect of secondary oxidants from VL~~ VL photo-oxidation under N_2 and air-saturated conditions

Lines 188–240: ~~As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions), while ^{3}VL* is the only oxidant expected under N_2-saturated conditions.~~ The photo-oxidation of VL ~~To examine the contributions of ^{3}VL*-derived secondary oxidants and ^{3}VL* only on VL photo-oxidation, experiments~~ under both N_2-~~air~~ and air-~~N_2~~-saturated conditions (Fig. S3a) were carried out at pH 4, which is representative of moderately acidic aerosol and cloud pH values (Pye et al., 2020). No significant VL loss was observed for dark experiments. The oxidation of ground-state VL by ^{3}VL* via H-atom

abstraction or electron transfer can form phenoxy (which is in resonance with a carbon-centered cyclohexadienyl radical that has a longer lifetime) and ketyl radicals (Neumann et al., 1986a, 1986b; Anastasio et al., 1996). The coupling of phenoxy radicals or phenoxy and cyclohexadienyl radicals can form oligomers as observed for both N_2- and air-saturated experiments (see discussions later). However, the little decay of VL under N_2-saturated condition indicates that these radicals probably predominantly decayed via back-hydrogen transfer to regenerate VL (Lathioor et al., 1999). A possible explanation for this is the involvement of O_2 in the secondary steps of VL decay. For instance, a major fate of the ketyl radical is reaction with O_2 (Anastasio et al., 1996). In the absence of O_2, radical formation occurs, but the forward reaction of ketyl radical and O_2 is blocked, leading to the regeneration of VL as suggested by the minimal VL decay. Aside from potential inhibition of secondary oxidants generation (Chen et al., 2020), N_2 purging may have also hindered the secondary steps for VL decay.

~~The low decay rate for VL* under N_2-saturated conditions suggests a minimal role for ^{3}VL* in VL photo-oxidation.~~ Contrastingly, the VL* decay rate constant under air-saturated conditions was 4 times higher~~, revealing the importance of ^{3}VL*-derived secondary oxidants for photosensitized oxidation of VL~~. As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions). However, the photo-oxidation of VL in this study is likely mainly governed by ^{3}VL* and that these secondary oxidants have only minor participation. ~~Aside from $^{\bullet}OH$, $O_2^{\bullet-}$/$^{\bullet}HO_2$ and $^1O_2$ can also promote VL photo-oxidation (Kaur and Anastasio, 2018; Chen et al., 2020).~~ $^1O_2$ is also a potential oxidant for phenols (Herrmann et al., 2010; Minella et al., 2011; Smith et al., 2014), but $^1O_2$ reacts much faster (by ~60 times) with phenolate ions compared to neutral phenols (Tratnyek and Hoigne, 1991; Canonica et al., 1995; McNally et al., 2005). Under the pH values (pH 2.5 to 4) considered in this study, the amount of phenolate ion is negligible, so the reaction between VL and $^1O_2$ should be slow. Interestingly, however, $^1O_2$ has been shown to be important in the photo-oxidation of 4-ethylguaiacol (pK_a = 10.3) by ^{3}C* of 3,4-dimethoxybenzaldehyde (solution with pH of ~3) (Chen et al., 2020). Furthermore, while the irradiation of other phenolic compounds can produce H_2O_2, a precursor for $^{\bullet}OH$ (Anastasio et al., 1996), the amount of H_2O_2 is small. Based on this, only trace amounts of H_2O_2 were likely generated from VL* (Li et al., 2014) under-air saturated conditions, suggesting that contribution from $^{\bullet}OH$ was minor. Overall, these suggest that VL photo-oxidation in this study is driven by ^{3}VL*. Further study on the impact of O_2 on the reactive intermediates involved is required to understand the exact mechanisms occurring under air-saturated conditions. Nonetheless, the VL* decay trends clearly indicate that O_2 is important for efficient VL photo-oxidation ~~an efficient oxidant for unsaturated organic compounds and has a lifetime that is much longer than ^{3}C* (Chen et al., 2020)~~. ~~Similar to VL*, t~~The decay rate constant for VL+AN under air-saturated conditions was also higher ~~faster~~ (6.6 times) than N_2-saturated conditions~~,~~ which may c~~an~~ be due to ~~several~~ reactions facilitated by nitrate photolysis products ~~and the enhancement of N(III)-mediated photo-oxidation~~ that may have been enhanced in the presence of O_2 ~~as reported in early works~~ (Vione et al., 2005; Kim et al., 2014; Pang et al., 2019a). As shown later, more nitrogen-containing species were observed under air-saturated conditions. An example is enhanced VL nitration likely from increased $^{\bullet}NO_2$ formation such as from the reaction of $^{\bullet}OH$ and $O_2^{\bullet-}$ with NO_2^- (Reactions 4 and 5, respectively; Table 1) or the autoxidation of $^{\bullet}NO$ from NO_2^- photolysis (Reactions 6–9; Table 1) in aqueous solutions (Pang et al., 2019a).

~~Reactions involving $^{\bullet}HO_2/O_2^{\bullet-}$ which may originate from the photolysis of nitrate alone, likely from the production and subsequent photolysis of peroxynitrous acid (HOONO) (Reaction 10; Table 1) (Jung et al., 2017; Wang et al., 2021), or the reactions of ^{3}VL* in the presence of O_2, may have contributed as well. For instance, Wang et al. (2021) recently demonstrated that nitrate photolysis generates $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and $HONO_{(g)}$ in the presence of dissolved aliphatic organic matter (e.g., nonanoic acid, ethanol), with the enhanced $HONO_{(g)}$ production caused by secondary photochemistry between $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and photoproduced $NO_{x(aq)}$ (Reactions 11 and 12; Table 1), in agreement with Scharko et al. (2014). The significance of this increased HONO production is enhanced $^{\bullet}$OH formation (Reaction 13; Table 1). In addition, $^{\bullet}HO_2$ can react with $^{\bullet}$NO (Reaction 10; Table 1) from NO_2^- photolysis (Reaction 6; Table 1) to form HOONO, and eventually $^{\bullet}NO_2$ and $^{\bullet}$OH (Reaction 14; Table 1) (Pang et al., 2019a).~~ Nevertheless, the comparable decay rate~~s~~ constants for VL* and VL+AN imply that ^{3}VL* chemistry still dominates even at 1:10 molar ratio of VL/nitrate, probably due to the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. Although we have no quantification of the oxidants in our reaction systems as it is outside the scope of this study, these observations clearly substantiate that~~secondary oxidants from ^{3}VL*, which are formed when O_2 is present, are required for efficient~~ photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation are more efficient in the presence of O_2.

Moreover, Section 3.1.3 (Effect of VOCs and inorganic anions) and related sentences have also been deleted based on the likely minor contribution of $^{\bullet}$OH to VL photo-oxidation in this study as pointed out by Reviewer 1 and suggested by other published literature (Anastasio et al., 1996; Li et al., 2014). Also, Section 3.5 has been deleted and Fig. 4, now 2, has been shown for the first time when potential aqSOA formation pathways were discussed, then referred to throughout the text. Table 1, which was reduced to half, was maintained in the revised version, while Table S2 (now 2) was moved to the main text, as suggested by the reviewer.

2. I do not trust using results for IPA to make generalizations about the effect of all VOCs on vanillin photooxidation. The authors repeat this questionable generalization several times throughout the manuscript, including twice in the abstract. Especially because the authors' explanation for the effect of IPA on their results relies on alcohol / water microstructure arguments, generalization to all VOCs seems unwarranted. Plus, IPA would be present only at very low concentrations in aqueous aerosol or cloud droplets due to its high volatility. It would be more appropriate if the authors remove (or heavily qualify) all statements about VOCs.

Response: We concur with the reviewer that the initial generalization for the effect of all VOCs on VL photo-oxidation based on IPA results is unsubstantiated. Moreover, the contribution of $^{\bullet}$OH to VL photo-oxidation in this study is likely minimal, as pointed out by Reviewer 1 and suggested by other published literature (Anastasio et al., 1996; Li et al., 2014). Section 3.1.3 (Effect of VOCs and inorganic anions) and related sentences have been deleted altogether. Nonetheless, we maintain that it would be worthwhile to explore the effects of other potential aerosol constituents on aqSOA formation and photo-oxidation studies (lines 627–631).

3. At several points, the authors discuss rather small differences between experiments (factors of 1.2 to 1.5) as significant, but the uncertainties in the parameter values being compared are never quantified. This raises doubts in readers' minds about which differences are actually statistically significant. Some discussion of uncertainties and random error is needed.

Response: Thank you for pointing this out. We have added relevant statements to discuss these uncertainties as follows:

Line 132: Each experiment was repeated independently at least three times and measurements were done in triplicate. The reported decay rate constants and absorbance enhancement are the average of results from triplicate experiments, and the corresponding errors represent one standard deviation.

The footnote of Table S2, now 2, has been revised as follows:

Table 2: [b]The data fitting was performed in the initial linear region. Each value is the average of results from triplicate experiments. Errors represent one standard deviation.

The revised text also now indicates whether a difference of less than a factor of 2 is statistically significant or not:

Line 331: The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$).

Line 510: The enhancement of GUA decay rate constant in the presence of VL is statistically significant ($p < 0.05$), while that in the presence of AN is not ($p > 0.05$).

4. The argument that ^{3}VL* is more reactive in its protonated form as an explanation for the observed pH effects does not make sense to me. The pKa of VL is 7.4, which means that more than 99.9% of it is protonated in all experiments, negating the possibility of any detectable acceleration at low pH by this mechanism. Furthermore, the authors describe reasonable alternative explanations for their observed pH effects, such as the more efficient photolysis of HONO vs NO2- producing more OH radicals at low pH. However, the questionable claim that 3VL* is more reactive in its protonated form is repeated several times throughout the manuscript (for example, lines 267, 270, 280, 449 and 500). This claim needs to be convincingly justified or removed from the manuscript.

Response: The pK_a (7.4) mentioned by the reviewer is for ground state VL, while the pK_a for the VL triplet has been reported to be 4.0 (Smith et al., 2016). As there are a greater fraction of VL triplets that are protonated at pH 2.5 (0.96) than at pH 4 (0.5), it is possible that the pH dependence observed in this study is due to ^{3}VL*being more reactive in its protonated form. Smith et al. (2016) also observed a pH dependence for the direct photodegradation of VL (0.005 mM) (rate constants at pH ≤ 3 are ~two times lower than at pH ≥ 5) which they attributed to the

sensitivity of the excimer of VL (i.e., the charge-transfer complex formed between an excited state VL molecule and a separate ground state VL molecule; Birks, 1973, Turro et al., 2010) to acid-base chemistry. The opposite trend observed in this study for 0.1 mM VL may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). It has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). We have added a statement to include the sensitivity of the excimer of VL to acid-base chemistry as another possibility for the observed pH dependence. In addition, we have removed the discussion of N(III) photolysis as an alternative explanation for the effects of pH on VL decay kinetics as there is no significant difference between the decay rate constants of VL* and VL+AN. Changes in the text are as follows:

Lines 331–352: The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$). The pK_a for the VL triplet has been reported to be 4.0 (Smith et al., 2016). As there are a greater fraction of VL triplets that are protonated at pH 2.5 (0.96) than at pH 4 (0.5), it is possible that the pH dependence of the decay rate constants observed in this study is due to ^{3}VL* being more reactive in its protonated form. Smith et al. (2016) also observed a pH dependence for the direct photodegradation of VL (0.005 mM) (rate constants at pH ≤ 3 are ~two times lower than at pH ≥ 5) which they attributed to the sensitivity of the excimer of VL (i.e., the charge-transfer complex formed between an excited state VL molecule and a separate ground state VL molecule; Birks, 1973, Turro et al., 2010) to acid-base chemistry. The opposite trend observed in this study for 0.1 mM VL may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). ~~For VL*, this pH trend indicates that ^{3}VL* are more reactive in their protonated form, which is opposite to that reported for 0.005 mM VL (Smith et al., 2016), likely due to the concentration dependence of the relative reactivities of protonated and neutral forms of ^{3}VL*.~~ It has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). ~~Also, increases in hydrogen ion concentration can enhance the formation of $HO_2^{\bullet}$ and H_2O_2 and in turn, $^{\bullet}$OH formation (Du et al., 2011). In addition to these pH influences on VL*, the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a) also contributed to the observed pH effects for VL+AN. At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}$OH formation than that of NO_2^- in the near-UV region (Arakaki et al., 1999; Kim et al., 2014). The increased $^{\bullet}$OH formation rates as pH decreases can lead to faster VL decay (Pang et al., 2019a). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}$OH (Reactions 4 and 15; Table 1) (Pang et al., 2019a).~~ As pH decreases, the higher reactivity of ^{3}VL* and sensitivity of the excimer of VL to acid-base chemistry ~~HONO being the dominant N(III) species can lead to faster VL photo-oxidation~~ may have led to faster VL photo-oxidation.

Line 356: further indicating that ^{3}VL* ~~are~~ may be more reactive in their protonated form.

Line 372: Essentially, the higher reactivity of ^{3}VL* and predominance of HONO over nitrite at lower pH may have resulted in increased formation of products mainly composed of oligomers and functionalized monomers.

Specific comments:

1. Line 25: The authors conclude that photosensitized reactions of VL were "more efficient" relative to nitrate-mediated photo-oxidation. However, as pointed out by the authors, VL is much more light-absorbing that nitrate. Can the authors make a comparative statement after taking this difference into account? Which is more efficient on a per-photon-absorbed basis? This would be a more appropriate comparison of reaction efficiency.

Response: Thank you for this suggestion. The apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of either VL or nitrate during simulated sunlight illumination can be defined as (Anastasio et al., 1996; Smith et al., 2014, 2016):

$$\Phi_{\mathrm{GUA}} = \frac{\text{mol GUA destroyed}}{\text{mol photons absorbed}} \qquad \text{(Eq. S9)}$$

Φ_{GUA} was calculated using the measured rate of GUA decay and rate of light absorption by either VL or nitrate through the following equation:

$$\Phi_{\mathrm{GUA}} = \frac{\text{rate of GUA decay}}{\text{rate of light absorption by VL or nitrate}} = \frac{k'_{\mathrm{GUA}} \times [\mathrm{GUA}]}{\sum[\left(1-10^{-\varepsilon_\lambda [\mathrm{C}] l}\right) \times I'_\lambda]} \qquad \text{(Eq. S10)}$$

where k'_{GUA} is the pseudo-first-order rate constant for GUA decay, [GUA] is the concentration of GUA (M), ε_λ is the base-10 molar absorptivity (M^{-1} cm^{-1}) of VL or nitrate at wavelength λ, [C] is the concentration of VL or nitrate (M), l is the pathlength of the illumination cell (cm), and I'_λ is the volume-averaged photon flux (mol-photons L^{-1} s^{-1} nm^{-1}) determined from 2NB actinometry:

$$j(\mathrm{2NB}) = 2.303 \times \Phi_{\mathrm{2NB}} \times l \times \Sigma_{300\,nm}^{350\,nm} (\varepsilon_{\mathrm{2NB},\lambda} \times I'_\lambda \times \Delta\lambda) \qquad \text{(Eq. S11)}$$

The ϕ_{GUA} in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL. We have revised this in the text as follows and added the information shown above in the supporting information: **Text S7**. Estimation of the apparent quantum efficiency of guaiacol photodegradation.

Line 29: Furthermore, comparisons of the apparent quantum efficiency of guaiacol photodegradation indicate that in this study, guaiacol oxidation by photosensitized reactions of VL ~~was observed to be more~~ is less efficient relative to nitrate-mediated photo-oxidation.

Other relevant revisions in the text are as follows:

Line 514: As mentioned earlier, ~~the~~ ^{3}VL* chemistry appears to be more important than that of nitrate photolysis even at 1:10 molar ratio of VL/nitrate on account of the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. However, the apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL (see Text S7 for the more details).

Line 619: The oxidation of guaiacol, a non-carbonyl phenol, by photosensitized reactions of vanillin was also shown to be ~~more~~ less efficient than that by nitrate photolysis products based on its lower apparent quantum efficiency.

2. Line 226: The authors at several points claim that VL triplet states and nitrate photolysis products have a "synergistic effect," but evidence in support of this claim is lacking, or at best the evidence supporting it is not adequately explained. The inadequately supported claim is repeated in line 497.

Response: Thank you for this point. The trends in line 297 (formerly 226) pertain to nitrate enhancing the increased normalized abundance of products and formation of more oxidized aqSOA from VL photo-oxidation in the presence of O_2 (VL* and VL+AN under air-saturated conditions) at pH 4, suggesting a potential enhancement of VL nitration in the presence of O_2. This has been revised as follows:

Lines 295–300: In brief, the presence of ~~secondary oxidants from ^{3}VL* and~~ O_2 increased the normalized abundance of products and promoted the formation of more oxidized aqSOA. ~~These trends were reinforced in the presence of nitrate, indicating synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ Compared to N_2-saturated condition, the higher normalized abundance of nitrogen-containing products under air-saturated condition for VL+AN (at pH 4) suggests a potential enhancement of VL nitration in the presence of O_2.

Line 606–611: ~~Further enhancement of VL photo-oxidation under air-saturated conditions in the presence of nitrate indicates synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~ In contrast, ^{3}VL*-initiated reactions proceeded rapidly under air-saturated conditions (O_2 is present) as indicated by higher VL decay rate constant and increased normalized abundance of products. For pH 4 experiments, the presence of both O_2 and nitrate resulted in the highest normalized abundance of products (including N-containing compounds) and $\langle OS_c \rangle$, which may be due to the presence of O_2 promoting VL nitration.

3. Line 258: This explanation of opposite pH trends at 0.1 and 0.005 mM VL is extremely speculative.

Response: Smith et al. (2016) reported that the relative reactivities of the protonated and neutral forms of VL triplets depend on the VL concentration. Specifically, they found that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). The opposite trend observed in our study may then be due to this concentration dependence of the reactivities of the protonated and neutral forms of ^{3}VL*. Moreover, as mentioned in our response to major comment #6 by Reviewer 1, the comparable pH dependence of the aqSOA formed from VL* at pH 4 and 2.5 over a range of pH conditions from 1.5 to 10.5 suggests that the pH dependence observed is likely due to the acid-base chemistry of the reactions which may involve ^{3}VL* or the excimer of VL (Smith et al., 2016). Please see our response to your major comment #4 for relevant changes made in the text.

4. Line 272: For greater clarity, it would be helpful if the manuscript would always match product formulas mentioned in the text to the structures shown in Table S3. Is this product structure #21 in Table S3?

Response: We thank the reviewer for this suggestion. We have added the product structures (if applicable) to the formulas mentioned in the text. No, product structure #21, now #12, refers to a GUA tetramer that was observed only in the GUA+VL experiment (Line 544). Unfortunately, we do not have a product structure for the tetramer mentioned in Line 358 (formerly 272), although we have now added the proposed formula for this tetramer as follows:

Line 358: Furthermore, a tetramer ($C_{31}H_{24}O_{11}$) was observed only in VL* at pH 2.5.

5. Line 297: is this dimer product structure #5 in Table S3?

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

6. Line 334: The solvent cage effect explanation seems questionable. Why would two negatively charged ions share a solvent cage, given their electrostatic repulsion? Furthermore, in line 339 the authors state that "NaBC did not cause any substantial change in the decay of VL," thus making this whole solvent cage discussion irrelevant to the data at hand.

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

7. Line 341 – 346: the authors state that "no tetramers were observed in VL*+NaBC" and "VL+AN+IPA had more oligomers," and then go on to suggest that the formation of oligomers can be promoted by inorganic ions, likely via the generation of radicals such as .CO3. No evidence has been provided, as far as I can tell, that NaBC promotes oligomer formation, so I was confused by the authors' claim here that bicarbonate does in fact promote oligomer formation via .CO3 radicals.

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

8. Line 363: ESI-MS is routinely used to detect macromolecules in biochemistry. This suggestion that the method cannot detect molecules with more than 25 carbons is an erroneous conclusion to draw from Lin et al. (2018).

Response: Lin et al. (2018) also studied BrC from BBOA, similar to our experiments. As they reported that majority of ESI-detected compounds in their study are smaller molecules with fewer than 25 carbon atoms, we raise the possibility that this may also be the case for our study.

9. Line 379: The logic needs to be better spelled out here. Why is the formation of more oxidized products suggested by a larger fraction of small-mass products observed for 1:1 VL/nitrate mixtures compared to 1:100? Do small product masses imply fragmentation, or is there a competition with oligomerization?

Response: We apologize for the confusion. The larger fraction of small product masses (< 200 m/z) observed for 1:1 compared to 1:100 VL/nitrate suggests the prevalence of functionalization for the former. In addition, the higher $<OS_c>$ for 1:1 VL/nitrate indicates the formation of more oxidized products compared to 1:100 VL/nitrate. The small product masses (< 200 m/z) imply functionalization, while the contribution of > 300 m/z suggests oligomerization. Fragmentation was indicated by the observed small organic acids (analyzed using IC), but not the mass spectrometric analyses as the small organic acids are not detectable in the positive ion mode. Fragmentation, which leads to the decomposition of previously formed oligomers and generation of small, oxygenated products such as organic acids (Huang et al., 2018), may also occur for the low [VL] experiments. However, its importance would likely be observed at longer irradiation times, similar to the high [VL] experiments. Oligomerization probably occurred for both 1:1 and 1:100 VL/nitrate as suggested by the observed contribution of > 300 m/z, although to a lesser extent than functionalization based on the higher contribution of < 200 m/z. Lines 466–485 have been revised as follows:

Lines 466–485: For both VL* and VL+AN, the contribution of < 200 m/z to the normalized abundance of products was higher at low [VL] than at high [VL], while the opposite was observed for > 300 m/z (Fig. S12~~3~~d). This indicates that functionalization was favored at low [VL], as supported by the higher $<OS_c>$, while oligomerization was the dominant pathway at high [VL], consistent with more oligomers or polymeric products reported from high phenols concentration (e.g., 0.1 to 3 mM) (Li et al., 2014; Slikboer et al., 2015; Ye et al., 2019). This is probably due to an increased concentration of phenoxy radicals (in resonance with a carbon-centered cyclohexadienyl radical) at high [VL], promoting radical-radical polymerization (Sun et al., 2010; Li et al., 2014). ~~Moreover~~ At low [VL], the contribution of < 200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, ~~further~~ suggesting the prevalence of functionalization for the former ~~formation of more oxidized products~~. ~~This may also be the reason why~~ In addition, 1:1 VL/nitrate (A11~~5~~; Table ~~S~~2) had higher $<OS_c>$ than 1:100 VL/nitrate (A12~~6~~; Table ~~S~~2), indicating the formation of more oxidized products, but had fewer N-containing compounds compared to the latter. A possible explanation is that at 1:1 VL/nitrate, VL may ~~efficiently~~ compete~~s~~ with NO_2^- for $^{\bullet}OH$ (from nitrate or nitrite photolysis, Reaction 4; Table 1) and indirectly reduce~~s~~ $^{\bullet}NO_2$. Similarly, hydroxylation has been suggested to be a~~n~~ more

important pathway for 1:1 VL/nitrite than in 1:10 VL/nitrite (Pang et al., 2019a). ~~This may also be the reason why 1:1 VL/nitrate (A15; Table S2) had higher <OS$_c$> than 1:100 (A16; Table S2) VL/nitrate but had fewer N-containing compounds compared to the latter.Moreover, the contribution of <200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, further suggesting the formation of more oxidized products.~~ Fragmentation, which leads to the decomposition of previously formed oligomers and generation of small, oxygenated products such as organic acids (Huang et al., 2018), may also occur for the low [VL] experiments. However, its importance would likely be observed at longer irradiation times, similar to the high [VL] experiments.

10. Line 389: $C_8H_9NO_3$ should be identified as product structure #2 (an amine) on Table S3.

Response: Thank you for this suggestion. We have now added the structure number to $C_8H_9NO_3$ in line 494 (formerly 389) as follows:

Line 494: … $C_8H_9NO_3$ (No. 2, Table S2) was also…

11. Line 408: The nitrate photolysis explanations may not be needed, given that the observed enhancement of nitrate on guaiacol decay rates was only a factor of 1.2. Is this a statistically significant change?

Response: Thank you for pointing this out. The reviewer is correct that the nitrate photolysis explanation is not needed in this case, given that the observed enhancement for GUA+AN is not statistically significant. Changes in the text are as follows:

Line 510: The enhancement of GUA decay rate constant in the presence of VL is statistically significant ($p < 0.05$), while that in the presence of AN is not ($p > 0.05$).

12. Line 418: The word "Similarly" is being used to relate two seemingly dissimilar observations, causing needless confusion. In the previous sentence, VL shows much higher absorbance enhancement than nitrate, but in this sentence nitrate is being compared to an experiment without nitrate.

Response: The reviewer is correct. The word 'similarly' has been removed in Line 525 (formerly 418) to avoid confusion.

Line 525: ~~Similarly,~~ Yang et al. (2021) also observed greater light absorption during nitrate-mediated photo-oxidation relative to direct GUA photodegradation.

13. Line 471: This sentence is confusing. Doesn't this work address (among other things) the effects of nitration on triplet-generating aromatics?

Response: Thank you for catching this. We have clarified this sentence as follows:

Line 574: Although nitrate did not substantially affect the VL decay rate~~s~~ constants~~,~~ likely due to much higher molar absorptivity of VL than nitrate and high VL concentration used in this work, the presence of nitrate promoted functionalization and nitration, indicating the significance of nitrate photolysis in this aqSOA formation pathway. ~~While~~ This work demonstrates that nitration, which is ~~can be~~ an important process for producing light-absorbing organics or BrC (Jacobson, 1999; Kahnt et al., 2013; Mohr et al., 2013; Laskin et al., 2015; Teich et al., 2017; Li et al., 2020), ~~its effect on~~ can also affect the aqueous-phase processing of triplet-generating aromatics ~~has not yet been examined in detail~~.

14. Line 481: Why would VL photodegrade 10 times slower in ALW relative to dilute cloudwater? This effect is important for applying this work to the atmosphere. Could the authors provide some theory or explanation here?

Response: There is no generalization yet for these increased or decreased photodegradation of methoxyphenols and by far, only VL has been observed to exhibit decreased photodegradation in ALW. The study for VL in ALW (Loisel et al., 2021) stated that the nature of inorganic ions would affect the photodegradation of organic compounds. Further work on the effects of inorganic ions on photodegradation of VL in ALW is warranted.

15. (a) On Table S2, experiments without nitrate are listed as "—" in the column of normalized abundances of N-containing compounds. Is this because no N-containing compounds were detected in the top 50, or because these samples were not analyzed for N-containing compounds? (b) It would be helpful to map the reactant molecule onto the Figure S12 graph.

Response: (a) The reviewer is right. The samples for experiments without nitrate were not analyzed for N-containing compounds. This information has been added to the footnote of Table S2 as follows:

…[d]The normalized abundance of products was calculated using Eq. 2. The samples for experiments without nitrate (marked with N/A ~~"—"~~) were not analyzed for N-containing compounds.

(b) VL has now been added to Fig. S12, now S11.

Technical Corrections:

1. Line 349: "increased" should be "increase"

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

2. Line 377: "an important" should be "a more important"

Response: Thank you for pointing this out. Line 478 (formerly 377) has been revised accordingly:

Line 478: Similarly, hydroxylation has been suggested to be a~~n~~ more important pathway for 1:1 VL/nitrite than in 1:10 VL/nitrite (Pang et al., 2019a).

3. Line 459: “decompose” should be “decomposes”

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

4. Sodium nitrate in my opinion would be better abbreviated “NaN” to be more consistent with other abbreviations such as “NaBC.”

Response: Sentences related to $^{\bullet}$OH scavengers have been deleted from the original manuscript.

5. Table S3: Compound number 4, the most abundant product in some studies, is missing an oxygen atom. It should be clarified that structure #1 is the reactant molecule vanillin rather than a product.

Response: Thank you for the correction and suggestion. We have corrected these on Table S2, formerly S3.

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Author Response for "Aqueous SOA formation from the photo-oxidation of vanillin: Direct photosensitized reactions and nitrate-mediated reactions" by Mabato et al.

We thank the reviewer for the thorough review and many constructive comments that helped improve the manuscript. Our point-by-point responses are below (changes to the original manuscript text and supporting information are in red, moved content in double-line strikethrough, and removed content in strikethrough). Please note that the line numbers in the responses refer to our revised manuscript with tracked changes. Also, please note that because we restructured the manuscript, the numbering of some figures and tables in the revised manuscript is different from those in the original manuscript.

Reviewer 3

This study investigated the aqueous photo-oxidation of vanillin (VL) via both direct photosensitized reaction and nitrate-mediated photo-oxidation and discussed the influence of secondary oxidants from triplet excited states (^{3}VL*), solution pH, VOCs, and inorganic anions, etc. in detail. The experiments and data analysis are well done, and the mechanisms that are proposed are plausible. This study provides valuable information about the chemical composition, optical properties, and possible reaction mechanisms for SOA formed from the VL photo-oxidation under different conditions. However, there are a few major and minor comments I would like the authors to address before it is considered for publication in ACP.

Major comments

(1) With the experiment design, it is difficult to directly compare ^{3}VL* pathway and nitrated-mediated pathway, as also mentioned by the authors that the VL concentration was very high, and ^{3}VL* chemistry dominated in all the VL + ammonium nitrate (AN) experiments. Maybe more precisely, what was compared was photo-oxidation of VL via ^{3}VL* chemistry with and without nitrate. However, both the title and some places in the manuscript are misleading.

Response: We agree with the reviewer that ^{3}VL* chemistry dominated in VL+AN experiments, as mentioned throughout the manuscript. However, we did not intend to compare ^{3}VL* and nitrated-mediated pathways either. It is our opinion that the current title does not have such connotation. Despite the high VL concentration, the VL+AN experiments showed some differences in the aqSOA formed. Hence, we would like to keep the current title. With the revised abstract, we hope it will not confuse readers into expecting a comparison.

(2) I suggest the authors restructure the manuscript: on the one hand, to move part of the figures and tables from the SI to the manuscript, e.g. Table S2 and Figure S12, to make it easier to follow. On the other hand, to simplify the article by cutting some "maybe interesting" but not that important/related findings/discussions to make the main storyline clearer.

Response: Thank you for the suggestion. Speculative discussions have been removed from the revised text. First, Section 3.1.1 has been amended to emphasize the importance of VL triplets:

Line 187: ~~Effect of secondary oxidants from VL~~ VL photo-oxidation under N_2 and air-saturated conditions

Lines 188–240: ~~As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions), while ^{3}VL* is the only oxidant expected under N_2-saturated conditions.~~ The photo-oxidation of VL ~~To examine the contributions of ^{3}VL*-derived secondary oxidants and ^{3}VL* only on VL photo-oxidation, experiments~~ under both N_2-~~air-~~ and air-~~N_2-~~saturated conditions (Fig. S3a) were carried out at pH 4, which is representative of moderately acidic aerosol and cloud pH values (Pye et al., 2020). No significant VL loss was observed for dark experiments. The oxidation of ground-state VL by ^{3}VL* via H-atom abstraction or electron transfer can form phenoxy (which is in resonance with a carbon-centered cyclohexadienyl radical that has a longer lifetime) and ketyl radicals (Neumann et al., 1986a, 1986b; Anastasio et al., 1996). The coupling of phenoxy radicals or phenoxy and cyclohexadienyl radicals can form oligomers as observed for both N_2- and air-saturated experiments (see discussions later). However, the little decay of VL under N_2-saturated condition indicates that these radicals probably predominantly decayed via back-hydrogen transfer to regenerate VL (Lathioor et al., 1999). A possible explanation for this is the involvement of O_2 in the secondary steps of VL decay. For instance, a major fate of the ketyl radical is reaction with O_2 (Anastasio et al., 1996). In the absence of O_2, radical formation occurs, but the forward reaction of ketyl radical and O_2 is blocked, leading to the regeneration of VL as suggested by the minimal VL decay. Aside from potential inhibition of secondary oxidants generation (Chen et al., 2020), N_2 purging may have also hindered the secondary steps for VL decay.

~~The low decay rate for VL* under N_2-saturated conditions suggests a minimal role for ^{3}VL* in VL photo-oxidation.~~ Contrastingly, the VL* decay rate constant under air-saturated conditions was 4 times higher~~, revealing the importance of ^{3}VL*-derived secondary oxidants for photosensitized oxidation of VL~~. As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions). However, the photo-oxidation of VL in this study is likely mainly governed by ^{3}VL* and that these secondary oxidants have only minor participation. ~~Aside from $^{\bullet}OH$, $O_2^{\bullet-}$/$^{\bullet}HO_2$ and $^1O_2$ can also promote VL photo-oxidation (Kaur and Anastasio, 2018; Chen et al., 2020).~~ $^1O_2$ is also a potential oxidant for phenols (Herrmann et al., 2010; Minella et al., 2011; Smith et al., 2014), but $^1O_2$ reacts much faster (by ~60 times) with phenolate ions compared to neutral phenols (Tratnyek and Hoigne, 1991; Canonica et al., 1995; McNally et al., 2005). Under the pH values (pH 2.5 to 4) considered in this study, the amount of phenolate ion is negligible, so the reaction between VL and $^1O_2$ should be slow. Interestingly, however, $^1O_2$ has been shown to be important in the photo-oxidation of 4-ethylguaiacol (pK_a = 10.3) by ^{3}C* of 3,4-dimethoxybenzaldehyde (solution with pH of ~3) (Chen et al., 2020). Furthermore, while the irradiation of other phenolic compounds can produce H_2O_2, a precursor for $^{\bullet}OH$ (Anastasio et al., 1996), the amount of H_2O_2 is small. Based on this, only trace amounts of H_2O_2 were likely generated from VL* (Li et al., 2014) under-air saturated conditions, suggesting that contribution from $^{\bullet}OH$ was minor. Overall, these suggest that VL photo-oxidation

in this study is driven by ^{3}VL*. Further study on the impact of O_2 on the reactive intermediates involved is required to understand the exact mechanisms occurring under air-saturated conditions. Nonetheless, the VL* decay trends clearly indicate that O_2 is important for efficient VL photo-oxidation ~~an efficient oxidant for unsaturated organic compounds and has a lifetime that is much longer than ^{3}C* (Chen et al., 2020)~~. ~~Similar to VL*, t~~The decay rate constant for VL+AN under air-saturated conditions was also higher ~~faster~~ (6.6 times) than N_2-saturated conditions~~,~~ which may c~~an~~ be due to ~~several~~ reactions facilitated by nitrate photolysis products ~~and the enhancement of N(III) mediated photo-oxidation~~ that may have been enhanced in the presence of O_2 ~~as reported in early works~~ (Vione et al., 2005; Kim et al., 2014; Pang et al., 2019a). As shown later, more nitrogen-containing species were observed under air-saturated conditions. An example is enhanced VL nitration likely from increased $^{\bullet}NO_2$ formation such as from the reaction of $^{\bullet}OH$ and $O_2^{\bullet-}$ with NO_2^- (Reactions 4 and 5, respectively; Table 1) or the autoxidation of $^{\bullet}NO$ from NO_2^- photolysis (Reactions 6–9; Table 1) in aqueous solutions (Pang et al., 2019a). ~~Reactions involving $^{\bullet}HO_2/O_2^{\bullet-}$ which may originate from the photolysis of nitrate alone, likely from the production and subsequent photolysis of peroxynitrous acid (HOONO) (Reaction 10; Table 1) (Jung et al., 2017; Wang et al., 2021), or the reactions of ^{3}VL* in the presence of O_2, may have contributed as well. For instance, Wang et al. (2021) recently demonstrated that nitrate photolysis generates $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and $HONO_{(g)}$ in the presence of dissolved aliphatic organic matter (e.g., nonanoic acid, ethanol), with the enhanced $HONO_{(g)}$ production caused by secondary photochemistry between $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and photoproduced $NO_{x(aq)}$ (Reactions 11 and 12; Table 1), in agreement with Scharko et al. (2014). The significance of this increased HONO production is enhanced $^{\bullet}OH$ formation (Reaction 13; Table 1). In addition, $^{\bullet}HO_2$ can react with $^{\bullet}NO$ (Reaction 10; Table 1) from NO_2^- photolysis (Reaction 6; Table 1) to form HOONO, and eventually $^{\bullet}NO_2$ and $^{\bullet}OH$ (Reaction 14; Table 1) (Pang et al., 2019a).~~ Nevertheless, the comparable decay rate~~s~~ constants for VL* and VL+AN imply that ^{3}VL* chemistry still dominates even at 1:10 molar ratio of VL/nitrate, probably due to the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. Although we have no quantification of the oxidants in our reaction systems as it is outside the scope of this study, these observations clearly substantiate that~~secondary oxidants from ^{3}VL*, which are formed when O_2 is present, are required for efficient~~ photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation are more efficient in the presence of O_2.

Moreover, Section 3.1.3 (Effect of VOCs and inorganic anions) and related sentences have also been deleted based on the likely minor contribution of $^{\bullet}OH$ to VL photo-oxidation in this study as pointed out by Reviewer 1 and suggested by other published literature (Anastasio et al., 1996; Li et al., 2014). Also, Section 3.5 has been deleted and Fig. 4, now 2, has been shown for the first time when potential aqSOA formation pathways were discussed, then referred to throughout the text. Table 1, which was reduced to half, was maintained in the revised version, while Table S2 (now 2) was moved to the main text, as suggested by the reviewer.

(3) (a) It is very interesting to see the changes in optical properties, and their relation to the changes in chemical composition. However, I only see very general discussions about it (e.g. line 234-238 and line 282-289). It will be nice to discuss the specific compounds, possible chromophores, and to explain the changes in the optical properties.

(b) To explain the pH-dependency, the authors cited Pang et al. 2019a, which reported the pH-dependent light absorbance of nitrophenols. However, the dominating products in this study were those without N, different from those in Pang et al. 2019a. In addition, the chemical composition of SOA with pH 4 and pH < 4 are quite different, which could also lead to different functional group/chromophores, and changes in optical properties.

Response:

(a) Thank you for this suggestion. We agree that identifying the BrC chromophores would enrich the discussion for the changes in the optical properties. However, it is possible that the products detected using UHPLC-qToF-MS in positive ESI mode might not have contributed significantly to all products formed and hence may not be the primary contributors to the absorbance enhancement. In other words, the absorbance enhancement may not necessarily correlate directly with the products detected. Detailed characterization of specific chromophores is indeed interesting, but it is outside of the scope of this study. Instead, the changes in the optical properties for this study are based on the integrated area of the absorption spectra from 350 to 550 nm in order to consider the contributions of all potentially light-absorbing products.

(b) The reviewer is correct that the dominant products in this study do not contain nitrogen. We have revised this explanation based on the comparable pH dependence of the aqSOA formed from VL* at pH 2.5 and 4 over a range of pH conditions from 1.5 to 10.5. This suggests that the observed pH dependence is due to the acid-base chemistry of the reactions, which may involve ^{3}VL* or the excimer of VL (Smith et al., 2016). Changes in the text are as follows:

Lines 375–386: ~~The h~~~~H~~igher absorbance enhancement for both VL* and VL+AN (Fig. 3~~2~~b) was observed as pH increased ~~may be attributed to redshifts and increased visible light absorption of reaction products (Pang et al., 2019a~~). To determine whether the pH dependence is due to the acid-base chemistry of the products or of the reactions, we measured the pH dependence of the aqSOA formed from VL* at pH 4 and 2.5 over a range of pH conditions from 1.5 to 10.5 (Fig. S10). For both cases, the intensity of absorption at longer wavelengths significantly increased as the pH of the solutions was raised. Moreover, the comparable pH dependence of the two solutions suggest that the observed pH dependence may be attributed to the acid-base chemistry of the reactions, which may involve ^{3}VL* or the excimer of VL (Smith et al., 2016), as discussed earlier. ~~When a phenolic molecule deprotonates at higher pH, an ortho- or para- electron-withdrawing group, such as a nitro or aldehyde group, can attract a portion of the negative charge towards its oxygen atoms through induced and conjugated effects, leading to the extension of chromophore from the electron-donating group (e.g., -O^{-}) to the electron-withdrawing group via the aromatic ring (Carey, 2000; Williams and Fleming, 2008; Pang et al., 2019a). Hence, the delocalization of the negative charge in phenolates leads to significant redshifts (Mohr et al., 2013).~~

For reference, changes in the revised text that are relevant to lines 375–386 are shown below:

Lines 331–352: The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$). The pK_a for the VL triplet has been reported to be 4.0 (Smith et al., 2016). As there are a greater fraction of VL triplets that are protonated at pH 2.5 (0.96) than at pH 4 (0.5), it is possible that the pH dependence of the decay rate constants observed in this study is due to ^{3}VL* being more reactive in its protonated form. Smith et al. (2016) also observed a pH dependence for the direct photodegradation of VL (0.005 mM) (rate constants at pH ≤ 3 are ~two times lower than at pH ≥ 5) which they attributed to the sensitivity of the excimer of VL (i.e., the charge-transfer complex formed between an excited state VL molecule and a separate ground state VL molecule; Birks, 1973, Turro et al., 2010) to acid-base chemistry. The opposite trend observed in this study for 0.1 mM VL may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). ~~For VL*, this pH trend indicates that ^{3}VL* are more reactive in their protonated form, which is opposite to that reported for 0.005 mM VL (Smith et al., 2016), likely due to the concentration dependence of the relative reactivities of protonated and neutral forms of ^{3}VL*.~~ It has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). ~~Also, increases in hydrogen ion concentration can enhance the formation of $HO_2^{\bullet}$ and H_2O_2 and in turn, $^{\bullet}$OH formation (Du et al., 2011). In addition to these pH influences on VL*, the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a) also contributed to the observed pH effects for VL+AN. At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}$OH formation than that of NO_2^- in the near-UV region (Arakaki et al., 1999; Kim et al., 2014). The increased $^{\bullet}$OH formation rates as pH decreases can lead to faster VL decay (Pang et al., 2019a). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}$OH (Reactions 4 and 15; Table 1) (Pang et al., 2019a).~~ As pH decreases, the higher reactivity of ^{3}VL* and sensitivity of the excimer of VL to acid-base chemistry ~~HONO being the dominant N(III) species can lead to faster VL photo-oxidation~~ may have led to faster VL photo-oxidation.

(4) Adding the experiments of guaiacol (GUA) is a little bit confusing, as the title is the photo-oxidation of VL. I understand it is a good addition to the manuscript, and these experiments nicely compared the photo-oxidation of GUV via the two pathways. However, the conclusion (line 25-26) "guaiacol oxidation by photosensitized reactions of VL was observed to be more efficient relative to nitrate-mediated photo-oxidation" is still problematic, as the concentration of VL in GUA + VL experiment was still 10 times higher than the observed value in the cloud and fog but the concentration of AN in GUA + AN experiments was similar to the observed concentration.

Response: Thank you for this suggestion. The apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of either VL or nitrate during simulated sunlight illumination can be defined as (Anastasio et al., 1996; Smith et al., 2014, 2016):

$$\Phi_{\mathrm{GUA}} = \frac{\text{mol GUA destroyed}}{\text{mol photons absorbed}} \quad \text{(Eq. S9)}$$

Φ_{GUA} was calculated using the measured rate of GUA decay and rate of light absorption by either VL or nitrate through the following equation:

$$\Phi_{\mathrm{GUA}} = \frac{\text{rate of GUA decay}}{\text{rate of light absorption by VL or nitrate}} = \frac{k'_{\mathrm{GUA}} \times [\mathrm{GUA}]}{\sum[\left(1-10^{-\varepsilon_{\lambda}[\mathrm{C}]l}\right) \times I'_{\lambda}\right)} \quad \text{(Eq. S10)}$$

where k'_{GUA} is the pseudo-first-order rate constant for GUA decay, [GUA] is the concentration of GUA (M), ε_{λ} is the base-10 molar absorptivity (M^{-1} cm^{-1}) of VL or nitrate at wavelength λ, [C] is the concentration of VL or nitrate (M), l is the pathlength of the illumination cell (cm), and I'_{λ} is the volume-averaged photon flux (mol-photons L^{-1} s^{-1} nm^{-1}) determined from 2NB actinometry:

$$j(\mathrm{2NB}) = 2.303 \times \Phi_{\mathrm{2NB}} \times l \times \Sigma_{300\,nm}^{350\,nm} (\varepsilon_{\mathrm{2NB},\lambda} \times I'_{\lambda} \times \Delta\lambda) \quad \text{(Eq. S11)}$$

The ϕ_{GUA} in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL. We have revised this in the text as follows and added the information shown above in the supporting information: **Text S7**. Estimation of the apparent quantum efficiency of guaiacol photodegradation.

Line 29: Furthermore, comparisons of the apparent quantum efficiency of guaiacol photodegradation indicate that in this study, guaiacol oxidation by photosensitized reactions of VL ~~was observed to be more~~ is less efficient relative to nitrate-mediated photo-oxidation.

Other relevant revisions in the text are as follows:

Line 514: As mentioned earlier, ~~the~~ ^{3}VL* chemistry appears to be more important than that of nitrate photolysis even at 1:10 molar ratio of VL/nitrate on account of the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. However, the apparent quantum efficiency of GUA photodegradation (ϕ_{GUA}) in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL (see Text S7 for the more details).

Line 619: The oxidation of guaiacol, a non-carbonyl phenol, by photosensitized reactions of vanillin was also shown to be ~~more~~ less efficient than that by nitrate photolysis products based on its lower apparent quantum efficiency.

Minor comments

1. Line 27-28 In the abstract, the sentence "which nitrate photolysis products can further enhance" sounds not clear to me.

Response: We apologize for the confusion. This sentence was supposed to convey two things: a) the direct photosensitized oxidation of VL may be an important aqSOA source and b) the addition of nitrate photolysis products to (a) can initiate further reactions that can enhance aqSOA formation. We have revised this sentence as follows:

Line 32: This study indicates that the direct photosensitized oxidation of VL~~,~~ and nitrate-mediated photo-oxidation of VL ~~which nitrate photolysis products can further enhance,~~ may be ~~an~~ important aqSOA sources in areas influenced by biomass burning emissions.

2. Line 121 Did you average these replicates for mass spectra and/or decay rates? Please clarify it.

Response: Yes, the reported mass spectra are based on the average of results from duplicate experiments, while the decay rate constants and absorbance enhancement are average of results from triplicate experiments. We have added the following sentence in the methods to clarify this:

Line 132: Each experiment was repeated independently at least three times and measurements were done in triplicate. The reported decay rate constants and absorbance enhancement are the average of results from triplicate experiments, and the corresponding errors represent one standard deviation.

The footnote of Table S2, now 2, has been revised as follows:

Table 2: [b]The data fitting was performed in the initial linear region. Each value is the average of results from triplicate experiments. Errors represent one standard deviation.

3. Line 168-169 It would be nice to explain it together with the chemical composition shown in Figure 1.

Response: This paragraph discusses the VL decay trends. Line 249 (formerly 196) already includes an example for this ($C_{16}H_{10}N_2O_9$; No. 3, Table S2). Changes in the text are as follows:

Lines 220: ~~Similar to VL*, t~~The decay rate constant for VL+AN under air-saturated conditions was also higher ~~faster~~ (6.6 times) than N_2-saturated conditions~~,~~ which may c~~an~~ be due to ~~several~~ reactions facilitated by nitrate photolysis products ~~and the enhancement of N(III)-mediated photo-oxidation~~ that may have been enhanced in the presence of O_2 ~~as reported in early works~~ (Vione et al., 2005; Kim et al., 2014; Pang et al., 2019a).

Lines 247–254: ~~Compared to N_2-saturated conditions, the normalized abundance of products such as o~~Oligomers, ~~and~~ functionalized monomers (e.g., demethylated VL; Fig. S4), and nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3~~3,~~ ~~Table S3~~,Table S2) (for VL+AN) had higher normalized abundance ~~were also more relatively abundant~~ under air-saturated conditions ~~were significantly higher under air-saturated conditions~~ (Figs. 1c-d), likely due to efficient ~~the secondary oxidants from~~ ^{3}VL*-initiated oxidation and enhanced VL nitration in the presence of ~~and~~ O_2 ~~and their interactions with nitrate photolysis products~~. ~~The nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S3) were also more relatively abundant under air-saturated conditions.~~ For both VL* and VL+AN under air-saturated conditions, the most abundant product was $C_{10}H_{10}O_5$ (No. 4, Table S2~~3~~), a substituted VL.

4. Line 181 "VL*" should be "^{3}VL*"?

Response: Thank you for catching this error. This has been corrected as follows:

Line 235: Nevertheless, the comparable decay rate~~s~~ constants for VL* and VL+AN imply that ^{3}VL* chemistry

5. Line 187 In both VL* and VL + AN under N2-saturated conditions (Fig. 1(a) and (b)), trimer signals are very high. Any explanations?

Response: N_2-saturated experiments would inhibit the formation of secondary oxidants, which can lead to ^{3}VL*-driven reactions (Chen et al., 2020) (line 107). Compared to $^{\bullet}$OH-mediated oxidation which yields more functionalized/oxygenated products, triplet-driven oxidation has been suggested to produce higher molecular weight products, probably with less fragmentation (Yu et al., 2014; Chen et al., 2020), as mentioned in lines 241–243. This likely explains the prevalence of dimers and trimers for the N_2-saturated experiments.

6. Line 212 Could you give some numbers to show "majority"?

Response: It is 58% of the 50 most abundant products for experiments A5 to A8. This has been added to the text as follows:

Lines 277–279: F~~f~~or experiments A5 to A8, ~~The~~ H:C ratios were mostly around 1.0 and double bond equivalent (DBE) values were typically (58% of the 50 most abundant products) > 7, indicating that the products ~~for experiments A5 to A8~~ ~~(Table S2)~~ were mainly oxidized aromatic ~~aromatic species~~ compounds (Xie et al., 2020).

7. Line 255 Should it be pH 4?

Response: Thank you for catching this error. The text has now been corrected as follows:

Line 331: the range of 2.5 to 4~~5~~

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Aqueous SOA formation from the photo-oxidation of vanillin: Direct photosensitized reactions and nitrate-mediated reactions

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Abstract. Vanillin (VL), a phenolic aromatic carbonyl abundant in biomass burning emissions, forms triplet excited states (^{3}VL*) under simulated sunlight leading to aqueous secondary organic aerosol (aqSOA) formation. This direct photosensitized oxidation of VL was compared with nitrate-mediated VL photo-oxidation under atmospherically relevant cloud and fog conditions~~,~~ through examining the VL decay kinetics, product compositions, and light absorbance changes. The majority of the most abundant products from both VL photo-oxidation pathways were potential Brown carbon (BrC) chromophores. In addition, both pathways generated oligomers, functionalized monomers, and oxygenated ring-opening products, but nitrate promoted functionalization and nitration, which can be ascribed to its photolysis products ($^{\bullet}$OH, $^{\bullet}NO_2$, and N(III), NO_2^- or HONO). Moreover, a potential imidazole derivative observed from nitrate-mediated VL photo-oxidation suggested that ammonium may be involved in the reactions. The effects of oxygen ~~(O_2)secondary oxidants from ^{3}VL*~~, pH, ~~the presence of volatile organic compounds (VOCs) and inorganic anions,~~ and reactants concentration and molar ratios on VL photo-oxidation were also explored. Our findings show that ~~the secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}$OH) from the reactions of ^{3}VL* and~~ O_2 plays an essential role in VL photo-oxidation, and ~~. Enhanced~~ oligomer formation was enhanced ~~noted~~ at pH < 4 ~~and in the presence of VOCs and inorganic anions, probably due to additional generation of radicals ($^{\bullet}HO_2$ and $CO_3^{\bullet-}$)~~. Also, functionalization was dominant at low VL concentration, whereas oligomerization was favored at high VL concentration. Furthermore, comparisons of the apparent quantum efficiency of guaiacol photodegradation indicate that in this study, guaiacol oxidation by photosensitized reactions of VL ~~was observed to be more~~ is less efficient relative to nitrate-mediated photo-oxidation. Lastly, potential aqSOA formation ~~VL photo-oxidation~~ pathways via VL photo-oxidation ~~under different reaction conditions~~ were proposed. This study indicates that the direct photosensitized oxidation of VL and nitrate-mediated VL photo-

oxidation ~~, which nitrate photolysis products can further enhance,~~ may be ~~an~~ important aqSOA sources in areas influenced by biomass burning emissions.

1 Introduction

Aqueous reactions can be an important source of secondary organic aerosols (SOA) (Blando and Turpin, 2000; Volkamer et al., 2009; Lim et al., 2010; Ervens et al., 2011; Huang et al., 2011; Lee et al., 2011; Smith et al., 2014) such as highly-oxygenated and low-volatility organics (Hoffmann et al., 2018; Liu et al., 2019) which may affect aerosol optical properties due to contributions to Brown Carbon (BrC) (Gilardoni et al., 2016). BrC refers to organic aerosols that absorb radiation efficiently in the near-ultraviolet (UV) and visible regions (Laskin et al., 2015). The formation of aqueous SOA (aqSOA) via photochemical reactions involves oxidation, with hydroxyl radical ($^{\bullet}OH$) usually considered as the primary oxidant (Herrmann et al., 2010; Smith et al., 2014). The significance of photosensitized chemistry in atmospheric aerosols has recently been reviewed (George et al., 2015). For instance, triplet excited states of organic compounds ($^3C^*$) from the irradiation of light-absorbing organics such as non-phenolic aromatic carbonyls (Canonica et al., 1995; Anastasio et al., 1996; Vione et al., 2006; Smith et al., 2014) have been reported to oxidize phenols at faster rates and with higher aqSOA yields compared to $^{\bullet}OH$ (Sun et al., 2010; Smith et al., 2014; Yu et al., 2014; Smith et al., 2016). Aside from being an oxidant, $^3C^*$ can also be a precursor of singlet oxygen ($^1O_2$), superoxide ($O_2^{\bullet-}$) or hydroperoxyl ($^{\bullet}HO_2$) radical, and $^{\bullet}OH$ (via $HO_2^{\bullet}/O_2^{\bullet-}$ formation) upon reactions with O_2 and substrates (e.g., phenols)~~, respectively~~ (~~Tinel~~George et al., 2018). The $^3C^*$ concentration in typical fog water has been estimated to be > 25 times than that of $^{\bullet}OH$, making $^3C^*$ the primary photo-oxidant for biomass burning phenolic compounds (Kaur and Anastasio, 2018; Kaur et al., 2019). Recent works on triplet-driven oxidation of phenols have mainly focused on changes of physicochemical properties (e.g., light absorption) and aqSOA yield (e.g., Smith et al., 2014, 2015, 2016), with few reports on reaction mechanisms and characterization of reaction products (e.g., Yu et al., 2014; Chen et al., 2020; Jiang et al., 2021).

Inorganic nitrate is a major component of aerosols and cloud/fog water. In cloud and fog water, the concentrations of inorganic nitrate can vary from 50 μM to > 1000 μM, with higher levels typically noted under polluted conditions (Munger et al., 1983; Collett et al., 1998; Zhang and Anastasio, 2003; Li et al., 2011; Giulianelli et al., 2014; Bianco et al., 2020). Upon photolysis (Vione et al., 2006; Herrmann, 2007; Scharko et al., 2014), inorganic nitrate in cloud and fog water can contribute to BrC (Minero et al., 2007) and aqSOA formation (Huang et al., 2018; Klodt et al., 2019; Zhang et al., 2021) by generating $^{\bullet}OH$ and nitrating agents (e.g., $^{\bullet}NO_2$). For example, the aqSOA yields from the photo-oxidation of phenolic carbonyls in nitrate are twice as high as that in sulfate solution (Huang et al., 2018). Nitration is a significant process in the formation of light-absorbing organics or BrC in the atmosphere (Jacobson, 1999; Kahnt et al., 2013; Mohr et al., 2013; Laskin et al., 2015; Teich et al., 2017; Li et al., 2020). Furthermore, nitrate photolysis has been proposed to be a potentially important process for SO_2

oxidation via the generation of $^{\bullet}OH$, $^{\bullet}NO_2$, and N(III) within particles (Gen et al., 2019a, 2019b), and it can also potentially change the morphology of atmospheric viscous particles (Liang et al., 2021). Accordingly, both $^3C^*$ and inorganic nitrate can contribute to aqSOA and BrC formation.

Biomass burning (BB) is a significant atmospheric source of both phenolic and non-phenolic aromatic carbonyls (Rogge et al., 1998; Nolte et al., 2001; Schauer et al., 2001; Bond et al., 2004). For ~~An~~ example, ~~is~~ vanillin (VL) (Henry's law constant of 4.56×10^5 M atm^{-1}; Yaws, 1994), a model compound for methoxyphenols which are abundant in BB emissions (Pang et al., 2019a), ~~which~~ has been shown to yield low-volatility products (Li et al., 2014) via aqueous $^{\bullet}OH$ oxidation and direct photodegradation. ~~–~~Photodegradation kinetics and aqSOA yields have been reported for direct VL photodegradation (Smith et al., 2016), with oxygenated aliphatic-like compounds (high H:C, $\geq$ 1.5 and low O:C, $\leq$ 0.5 ratios) reported as the most likely products (Loisel et al., 2021). Additionally, aqueous-phase reactions of phenols with reactive nitrogen species have been proposed to be a significant source of nitrophenols and SOA (Grosjean, 1985; Kitanovski et al., 2014; Kroflič et al., 2015; Pang et al., 2019a; Kroflič et al., 2021; Yang et al., 2021). For instance, nitrite-mediated VL photo-oxidation can generate nitrophenols, and the reactions are influenced by nitrite/VL molar ratios, pH, and the presence of $^{\bullet}OH$ scavengers (Pang et al., 2019a). As BB aerosols are typically internally mixed with other aerosol components (Zielinski et al., 2020), VL may coexist with nitrate in BB aerosols. The aqueous-phase photo-oxidation of VL and nitrate may then reveal insights into the atmospheric processing of BB aerosols.

~~As BB aerosols are typically internally mixed with other aerosol components (Zielinski et al., 2020), VL may coexist with nitrate in BB aerosols. The aqueous-phase photo-oxidation of VL and nitrate may then reveal insights into the atmospheric processing of BB aerosols. In addition, pollution from large BB events in central Amazonia has been reported to interact with volatile organic compounds (VOCs) and soil dust (Rizzo et al., 2010). Moreover, the production, growth, and chemical complexity of SOA can be influenced by the uptake and aerosol phase reactions of VOCs (Pöschl, 2005; De Gouw and Jimenez, 2009; Ziemann and Atkinson, 2012). Accordingly, studies incorporating other atmospherically relevant species (e.g., VOCs and inorganic anions) in photo-oxidation experiments are warranted.~~

To evaluate the potential significance of VL and its reactions with nitrate in aqSOA formation in cloud/fog water, we studied the direct photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation under atmospherically relevant conditions. In this work, reactions were characterized based on VL decay kinetics, light absorbance changes, and products. The influences of ~~-~~ O_2~~secondary oxidants from VL triplets~~, solution pH, ~~the presence of VOCs and inorganic anions,~~ and reactants concentration and molar ratios on these two photo-oxidation pathways were also assessed. The $^3C^*$ of non-phenolic aromatic carbonyls (e.g., 3-4-dimethoxybenzaldehyde, DMB; a non-phenolic aromatic carbonyl) (Smith et al., 2014; Yu et al., 2014; Jiang et al., 2021) and phenolic aromatic carbonyls (e.g., acetosyringone, vanillin) (Smith et al., 2016) have been shown to oxidize phenols, but the reaction products from the latter are unknown. We then examined the photo-oxidation of guaiacol, another non-carbonyl phenol, in the presence of VL and compared it with nitrate-mediated photo-oxidation. Finally, we proposed aqSOA formation pathways via VL photo-oxidation ~~pathways of VL under different reaction conditions~~. This work presents a comprehensive comparison of VL photo-oxidation by VL photosensitization and in the presence of inorganic nitrate.

2 Methods

2.1 Aqueous phase photo-oxidation experiments

Photo-oxidation experiments were performed in a ~~500 mL~~ custom-built quartz photoreactor. The solutions (initial volume of 500 mL) were continuously mixed throughout the experiments using ~~equipped with~~ a magnetic stirrer. The solutions were bubbled with synthetic air or nitrogen (N_2) (> 99.995%) (0.5 dm^3/min) for 30 min before irradiation to achieve air- and N_2-saturated conditions, respectively, and the bubbling was continued throughout the reactions (Du et al., 2011; Chen et al., 2020). The aim of the air-saturated experiments was to enable the generation of secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) from $^3VL^*$ as O_2 is present. Conversely, the N_2-saturated experiments would inhibit the formation of these secondary oxidants, which can lead~~ing~~ to $^3VL^*$-driven reactions (Chen et al., 2020). Solutions were irradiated through the quartz window of the reactor using a xenon lamp (model 6258, Ozone free xenon lamp, 300 W, Newport) equipped with a longpass filter (20CGA-305 nm cut-on filter, Newport) to eliminate light below 300 nm. Cooling fans positioned around the photoreactor and lamp housing maintained reaction temperatures at 27 ± 2 °C. The averaged initial photon flux in the reactor from 300 to 380 nm measured using a chemical actinometer (2-nitrobenzaldehyde) was 2.6×10^{15} photons cm^{-2} s^{-1} nm^{-1} (Fig. S1). Although the concentration of VL in cloud/fog water has been estimated to be < 0.01 mM (Anastasio et al., 1996), a higher VL concentration (0.1 mM) was used in this study to guarantee sufficient signals for product identification (Vione et al., 2019). The chosen ammonium nitrate (AN) concentration (1 mM) was based on values observed in cloud and fog water (Munger et al., 1983; Collett et al., 1998; Zhang and Anastasio, 2003; Li et al., 2011; Giulianelli et al., 2014; Bianco et al., 2020). It should be noted that this study is not intended to identify the concentrations of nitrate that would affect the kinetics. ~~We also examined the role of VOCs (2-propanol, IPA) (1 mM) and inorganic anions (sodium bicarbonate, NaBC) (1 mM) in these reactions. IPA can be classified as both a biogenic (from grass, Olofsson et al., 2003) and anthropogenic VOC (e.g., from solvents and industrial processes, Hippelein, 2004; Lewis et al., 2020), while bicarbonate is an inorganic anion observed in fog water from both urban and rural locations (Collett et al., 1999; Straub et al., 2012; Straub, 2017). IPA and NaBC are particularly interesting also because they can produce other radicals (e.g., $^{\bullet}HO_2$ and carbonate radical, $CO_3^{\bullet-}$) that may react with nitrate photolysis products (Vione et al., 2009; Wang et al., 2021) and they can act as $^{\bullet}OH$ scavengers (Warneck and Wurzinger, 1988; Vione et al., 2009; Gen et al., 2019b; Pang et al., 2019a), although it must be noted that these compounds were not added in excess for our experiments.~~ Moreover, comparisons were made between the photo-oxidation of guaiacol (0.1 mM), a non-carbonyl phenol, in the presence of VL (0.1 mM) or AN (1 mM). Samples (10 mL) were collected hourly for a total of 6 h for offline optical and chemical analyses. Absorbance measurements, VL (and GUA) decay kinetics (calibration curves for VL and GUA standard solutions; Fig. S2), small organic acids measurements, and product characterization were conducted using UV-Vis spectrophotometry, ultra-high-performance liquid chromatography with photodiode array detector (UHPLC-PDA), ion chromatography (IC), and UHPLC coupled with quadrupole time-of-flight mass spectrometry (UHPLC-qToF-MS) equipped with an electrospray

ionization (ESI) source and operated in the positive ion mode (the negative ion mode signals were too low for our analyses), respectively. Each experiment was repeated independently at least three times and measurements were done in triplicate. The reported decay rate constants and absorbance enhancement are the average of results from triplicate experiments, and the corresponding errors represent one standard deviation. The mass spectra are based on the average of results from duplicate experiments. ~~Details on the materials and analytical procedures are provided in the Supporting Information (Text S1 to S6)~~The Supporting Information (Text S1 to S7) provides details on the materials and analytical procedures. The pseudo-first-order rate constant (*k*') for VL decay was determined using the following equation (Huang et al., 2018):

$$ln\,([\mathrm{VL}]_t/[\mathrm{VL}]_0) = -k't \qquad \text{(Eq. 1)}$$

where $[\mathrm{VL}]_t$ and $[\mathrm{VL}]_0$ are the concentrations of VL at time *t* and 0, respectively. Replacing VL with GUA in Eq. 1 enabled the calculation of GUA decay rate constant. The decay rate constants were normalized to the photon flux measured for each experiment through dividing *k'* by the measured 2-nitrobenzaldehyde (2NB) decay rate constant, *j*(2NB) (see Text S6 for more details).

2.2 Calculation of normalized abundance of products

Comparisons of peak abundance in mass spectrometry have been used in many recent studies (e.g., Lee et al., 2014; Romonosky et al., 2017; Wang et al., 2017; Fleming et al., 2018; Song et al., 2018; Klodt et al., 2019; Ning et al., 2019) to show the relative importance of different types of compounds (Wang et al., 2021). However, ionization efficiency may greatly vary for different classes of compounds (Kebarle, 2000; Schmidt et al., 2006; Leito et al., 2008; Perry et al., 2008; Kruve et al., 2014)~~,~~ and so uncertainties may arise from comparisons of peak areas among compounds. In this work, we assumed equal ionization efficiency of different compounds, which is commonly used to estimate O:C ratios of SOA (Bateman et al., 2012; Lin et al., 2012; De Haan et al., 2019), to calculate their normalized abundance. The normalized abundance of a product, [P] (unitless), was calculated as follows:

$$[\mathrm{P}] = \frac{\mathrm{A}_{P,t}}{\mathrm{A}_{VL,t}} \cdot \frac{[\mathrm{VL}]_t}{[\mathrm{VL}]_0} \qquad \text{(Eq. 2)}$$

where $\mathrm{A}_{P,t}$ and $\mathrm{A}_{VL,t}$ are the extracted ion chromatogram (EIC) ~~signal~~ peak areas of the product P and VL from UHPLC-qToF-MS analyses at time *t*, respectively; $[\mathrm{VL}]_t$ and $[\mathrm{VL}]_0$ are the VL concentrations (μM) determined using UHPLC at time *t* and 0, respectively. Here, we relied on the direct quantification ~~more accurate measurements~~ of [VL] using UHPLC~~-~~(see Fig. S2 for VL calibration curve)~~for semi-quantification~~. ~~It should be noted that the ionization efficiency may greatly vary for different classes of compounds (Kebarle, 2000). Hence, we assumed equal ionization efficiency of different compounds to calculate their normalized abundance, which is commonly used to estimate O:C ratios of SOA (Bateman et al., 2012; Lin et al., 2012; De Haan et al., 2019).~~ It should be noted that the normalized abundance of products in this study is a semi-quantitative analysis intended to provide an overview of how the signal intensities changed under different experimental conditions but not to

quantify the absolute concentration of products. Moreover, the major products detected in this study are probably those with high concentration or high ionization efficiency in the positive ESI mode. The use of relative abundance (product peaks are normalized to the highest peak) (e.g., Lee et al., 2014; Romonosky et al., 2017; Fleming et al., 2018; Klodt et al., 2019) would yield the same major products reported. Typical fragmentation behavior observed in MS/MS spectra for individual functional groups from Holčapek et al. (2010) are outlined in Table S1.

3 Results and Discussion

3.1 Kinetics, mass spectrometric, and absorbance changes analyses during aqueous phase photo-oxidation of vanillin

For clarity purposes, the reactions involving reactive species referred to in the following discussions are provided in Table 1. Table ~~S~~2 summarizes the reaction conditions, initial VL (and GUA) decay rate~~s~~ constants, normalized abundance of products, and average carbon oxidation state (<OS_c> (of the 50 most abundant products). In general, the 50 most abundant products contributed more than half of the total normalized abundance of products. ~~For clarity purposes, the reactions involving reactive species referred to in the following discussions are provided in Table 1.~~

As shown in Figure S3, VL underwent oxidation both directly and in the presence of nitrate upon simulated sunlight illumination. VL absorbs light and is promoted to its excited singlet state ($^1VL^*$), then undergoes intersystem crossing (ISC) to the excited triplet state, $^3VL^*$. In principle, $^3VL^*$ can oxidize ground-state VL (Type I photosensitized reactions) via H-atom abstraction/electron transfer and form $O_2^{\bullet-}$ or $HO_2^{\bullet}$ in the presence of O_2 (~~Tinel~~George et al., 2018), or react with O_2 (Type II photosensitized reactions) to yield $^1O_2$ via energy transfer or $O_2^{\bullet-}$ via electron transfer (Lee et al., 1987; Foote et al., 1991). The disproportionation of $HO_2^{\bullet}/O_2^{\bullet-}$ (Anastasio et al., 1996) ~~and reaction of $HO_2^{\bullet}$ with $O_2^{\bullet-}$ (Du et al., 2011)~~ form hydrogen peroxide (H_2O_2), which is a photolytic source of $^{\bullet}OH$. Overall, air-saturated conditions, in which O_2 is present, enable the generation of secondary oxidants from $^3VL^*$ ($^1O_2$, $O_2^{\bullet-}/^{\bullet}HO_2$, $^{\bullet}OH$). Moreover, $^{\bullet}OH$, $^{\bullet}NO_2$, and NO_2^-/HNO_2, i.e., N(III), generated via nitrate photolysis (Reactions 1~~-~~–3; Table 1), can also oxidize or nitrate VL. In this work, the direct photosensitized oxidation of VL (~~(by $^3VL^*$ or secondary oxidants from $^3VL^*$ and O_2~~VL only experiments) and nitrate-mediated VL photo-oxidation are referred to as VL* and VL+AN, respectively.

3.1.1 ~~Effect of secondary oxidants from VL triplets~~VL photo-oxidation under N_2 and air-saturated conditions

~~As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}/^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from $^3VL^*$ when O_2 is present (e.g., under air-saturated conditions), while $^3VL^*$ is the only oxidant expected under N_2-saturated conditions.~~ The photo-oxidation of VL ~~To examine the contributions of $^3VL^*$-derived secondary oxidants and $^3VL^*$ only on VL photo-oxidation, experiments~~ under both N_2- ~~air~~ and air-~~N_2~~ saturated conditions (Fig. S3a) were carried out at pH 4, which is representative of moderately acidic aerosol and cloud pH values (Pye et al., 2020). No significant VL loss was observed for dark experiments. The oxidation of ground-state VL by $^3VL^*$ via H-atom abstraction or electron transfer can form phenoxy (which is in resonance with a carbon-

centered cyclohexadienyl radical that has a longer lifetime) and ketyl radicals (Neumann et al., 1986a, 1986b; Anastasio et al., 1996). The coupling of phenoxy radicals or phenoxy and cyclohexadienyl radicals can form oligomers as observed for both N_2- and air-saturated experiments (see discussions later). However, the little decay of VL under N_2-saturated condition indicates that these radicals probably predominantly decayed via back-hydrogen transfer to regenerate VL (Lathioor et al., 1999). A possible explanation for this is the involvement of O_2 in the secondary steps of VL decay. For instance, a major fate of the ketyl radical is reaction with O_2 (Anastasio et al., 1996). In the absence of O_2, radical formation occurs, but the forward reaction of ketyl radical and O_2 is blocked, leading to the regeneration of VL as suggested by the minimal VL decay. Aside from potential inhibition of secondary oxidants generation (Chen et al., 2020), N_2 purging may have also hindered the secondary steps for VL decay.

~~The low decay rate for VL* under N_2-saturated conditions suggests a minimal role for ^{3}VL* in VL photo-oxidation.~~ Contrastingly, the VL* decay rate constant under air-saturated conditions was 4 times higher~~, revealing the importance of ^{3}VL*-derived secondary oxidants for photosensitized oxidation of VL~~. As mentioned earlier, secondary oxidants ($^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) can be generated from ^{3}VL* when O_2 is present (e.g., under air-saturated conditions). However, the photo-oxidation of VL in this study is likely mainly governed by ^{3}VL* and that these secondary oxidants have only minor participation. ~~Aside from $^{\bullet}OH$, $O_2^{\bullet-}$/$^{\bullet}HO_2$ and $^1O_2$ can also promote VL photo-oxidation (Kaur and Anastasio, 2018; Chen et al., 2020).~~ $^1O_2$ is also a potential oxidant for phenols (Herrmann et al., 2010; Minella et al., 2011; Smith et al., 2014), but $^1O_2$ reacts much faster (by ~60 times) with phenolate ions compared to neutral phenols (Tratnyek and Hoigne, 1991; Canonica et al., 1995; McNally et al., 2005). Under the pH values (pH 2.5 to 4) considered in this study, the amount of phenolate ion is negligible, so the reaction between VL and $^1O_2$ should be slow. Interestingly, however, $^1O_2$ has been shown to be important in the photo-oxidation of 4-ethylguaiacol (pK_a = 10.3) by ^{3}C* of 3,4-dimethoxybenzaldehyde (solution with pH of ~3) (Chen et al., 2020). Furthermore, while the irradiation of other phenolic compounds can produce H_2O_2, a precursor for $^{\bullet}OH$ (Anastasio et al., 1996), the amount of H_2O_2 is small. Based on this, only trace amounts of H_2O_2 were likely generated from VL* (Li et al., 2014) under-air saturated conditions, suggesting that contribution from $^{\bullet}OH$ was minor. Overall, these suggest that VL photo-oxidation in this study is driven by ^{3}VL*. Further study on the impact of O_2 on the reactive intermediates involved is required to understand the exact mechanisms occurring under air-saturated conditions. Nonetheless, the VL* decay trends clearly indicate that O_2 is important for efficient VL photo-oxidation~~an efficient oxidant for unsaturated organic compounds and has a lifetime that is much longer than ^{3}C* (Chen et al., 2020)~~. ~~Similar to VL*, t~~The decay rate constant for VL+AN under air-saturated conditions was also higher ~~faster~~ (6.6 times) than N_2-saturated conditions, which may ~~can~~ be due to ~~several~~ reactions facilitated by nitrate photolysis products ~~and the enhancement of N(III) mediated photo-oxidation~~ that may have been enhanced in the presence of O_2 ~~as reported in early works~~ (Vione et al., 2005; Kim et al., 2014; Pang et al., 2019a). As shown later, more nitrogen-containing species were observed under air-saturated conditions. An example is enhanced VL nitration likely from increased $^{\bullet}NO_2$ formation such as from the reaction of $^{\bullet}OH$ and $O_2^{\bullet-}$ with NO_2^- (Reactions 4 and 5, respectively; Table 1) or the autoxidation of $^{\bullet}NO$ from NO_2^- photolysis (Reactions 6–~~-~~9; Table 1) in aqueous solutions (Pang

et al., 2019a). ~~Reactions involving $^{\bullet}HO_2/O_2^{\bullet-}$ which may originate from the photolysis of nitrate alone, likely from the production and subsequent photolysis of peroxynitrous acid (HOONO) (Reaction 10; Table 1) (Jung et al., 2017; Wang et al., 2021), or the reactions of ^{3}VL* in the presence of O_2, may have contributed as well. For instance, Wang et al. (2021) recently demonstrated that nitrate photolysis generates $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and $HONO_{(g)}$ in the presence of dissolved aliphatic organic matter (e.g., nonanoic acid, ethanol), with the enhanced $HONO_{(g)}$ production caused by secondary photochemistry between $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and photoproduced $NO_{x(aq)}$ (Reactions 11 and 12; Table 1), in agreement with Scharko et al. (2014). The significance of this increased HONO production is enhanced $^{\bullet}$OH formation (Reaction 13; Table 1). In addition, $^{\bullet}HO_2$ can react with $^{\bullet}$NO (Reaction 10; Table 1) from NO_2^- photolysis (Reaction 6; Table 1) to form HOONO, and eventually $^{\bullet}NO_2$ and $^{\bullet}$OH (Reaction 14; Table 1) (Pang et al., 2019a).~~ Nevertheless, the comparable decay rate~~s~~ constants for VL* and VL+AN imply that ^{3}VL* chemistry still dominates even at 1:10 molar ratio of VL/nitrate, probably due to the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. Although we have no quantification of the oxidants in our reaction systems as it is outside the scope of this study, these observations clearly substantiate that ~~secondary oxidants from ^{3}VL*, which are formed when O_2 is present, are required for efficient~~ photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation are more efficient in the presence of O_2.

The products from VL* under N_2-saturated conditions were mainly oligomers (e.g., $C_{16}H_{14}O_4$) (Fig. 1a), consistent with triplet-mediated oxidation forming higher molecular weight products, probably with less fragmentation relative to oxidation by $^{\bullet}$OH (Yu et al., 2014; Chen et al., 2020). A threefold increase in the normalized abundance of products was noted upon addition of nitrate (VL+AN under N_2-saturated conditions; Fig. 1b), and in addition to oligomers, functionalized monomers (e.g., ~~-~~$C_8H_6O_5$) and nitrogen-containing compounds (e.g., $C_8H_9NO_3$; No. 2, Table S2~~3~~) were also observed, in agreement with $^{\bullet}$OH-initiated oxidation yielding more functionalized/oxygenated products compared to triplet-mediated oxidation (Yu et al., 2014; Chen et al., 2020). ~~Compared to N_2-saturated conditions, the normalized abundance of products such as o~~Oligomers, ~~and~~ functionalized monomers (e.g., demethylated VL; Fig. S4), and ~~The~~ nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3~~3, Table S3~~, Table S2) (for VL+AN) had higher normalized abundance~~were also more relatively abundant~~ under air-saturated conditions ~~were significantly higher under air saturated conditions~~ (Figs. 1c-d), likely due to efficient ~~the secondary oxidants from~~ ^{3}VL*-initiated oxidation and enhanced VL nitration in the presence of ~~and~~ O_2 ~~and their interactions with nitrate photolysis products~~. ~~The nitrogen-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S3) were also more relatively abundant under air-saturated conditions.~~ For both VL* and VL+AN under air-saturated conditions, the most abundant product was $C_{10}H_{10}O_5$ (No. 4, Table S2~~3~~), a substituted VL.~~-~~ Irradiation of VL by 254-nm lamp has also been reported to lead to VL dimerization and functionalization via ring-retaining pathways, as well as small oxygenates but only when $^{\bullet}$OH from H_2O_2 were involved (Li et al., 2014). In this work, small organic acids (e.g., formic acid) were observed from both VL* and VL+AN under air-saturated conditions (Fig. S5) due to simulated sunlight that could access the 308-nm VL band (Smith et al., 2016). Interestingly, we observed a potential imidazole derivative ($C_5H_5N_3O_2$; ~~Fig. 1d~~No. 5, Table S2) from VL+AN under air-saturated conditions (Fig. 1d), which may have formed from reactions induced by ammonium. This compound was not observed in a parallel experiment in which AN was replaced with sodium nitrate (SN) (Fig. S6a; see Sect.

3.3 for discussion). The potential aqSOA formation ~~most probable~~ pathways via ~~of direct photosensitized and nitrate-mediated photo-oxidation of~~ VL photo-oxidation in this study are summarized in Fig. 2~~were proposed (Fig. 4). In Scheme 1 (pH 4 and pH <4 under air-saturated conditions), ^{3}VL* and $^{\bullet}$OH (from ^{3}VL* or nitrate photolysis) can initiate H-atom abstraction to generate phenoxy or ketyl radicals (Huang et al., 2018; Vione et al., 2019).~~ At pH 4, ring-opening products (Fig. S5) from fragmentation in both VL* and VL+AN may have reacted with VL or dissolved ammonia to generate $C_{10}H_{10}O_5$ (No. 4, Table S2) (Pang et al., 2019b) or~~and~~ a potential imidazole derivative ($C_5H_5N_3O_2$; No. 5, Table S2), respectively. Moreover, nitrate photolysis products promoted functionalization and nitration (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S2). ~~At pH <4, the reactivity of ^{3}VL* increased as suggested by the abundance of oligomers (e.g., $C_{16}H_{14}O_6$) and increased normalized abundance of N-containing compounds.~~

The molecular transformation of VL upon photo-oxidation was examined using the van Krevelen diagrams (Fig. S7). For all experiments (A1-15~~9~~; Table ~~S~~2) in this study, the O:C and H:C ratios of the products were mainly similar to those observed from the aging of other phenolics (Yu et al., 2014) and BB aerosols (Qi et al., 2019). U~~u~~nder N_2-saturated conditions, o~~O~~ligomers with O:C ratios ≤ 0.6 were dominant in VL*, ~~under N_2-saturated conditions.~~ while smaller molecules (n_c ≤ 8) with higher O:C ratios (up to 0.8) were also observed f~~F~~or VL+AN. ~~For VL+AN under N_2-saturated conditions, smaller molecules (n_c ≤8) with higher O:C ratios (up to 0.8) were also observed.~~ In contrast, ~~M~~more products with higher O:C ratios (≥ 0.6) were noted under air-saturated conditions for both VL* and VL+AN. F~~f~~or experiments A5 to A8, ~~The~~ H:C ratios were mostly around 1.0 and double bond equivalent (DBE) values were typically (58% of the 50 most abundant products) ≥ 7, indicating that the products ~~for experiments A5 to A8 (Table S2)~~ were mainly oxidized aromatic ~~aromatic species~~ compounds (Xie et al., 2020). Compounds with H:C ≤ 1.0 and O:C ≤ 0.5 are common for aromatic species, while compounds with H:C ≥ 1.5 and O:C ≤ 0.5 are typical for more aliphatic species (Mazzoleni et al., 2012; Kourtchev et al., 2014; Jiang et al., 2021). ~~Moreover, majority of the products for experiments A5 to A8 have double bond equivalent (DBE) values >7, which corresponds to oxidized aromatic compounds (Xie et al., 2020).~~ In contrast, Loisel et al. (2021) reported mainly oxygenated aliphatic-like compounds ~~(H:C, ≥1.5 and O:C, ≤0.5 ratios)~~ from the direct irradiation of VL (0.1 mM), ~~which may be~~probably due to their use of ESI in the negative ion mode, which has higher sensitivity for detecting compounds such as carboxylic acids (Holčapek et al., 2010; Liigand et al., 2017), and solid-phase extraction (SPE) procedure causing the loss of some oligomers (LeClair et al., 2012; Zhao et al., 2013; Bianco et al., 2018). Among experiments A5 to A8 ~~(Table S2)~~, VL+AN under air-saturated conditions (A7) had the highest normalized abundance of products and <OS_c>, ~~most~~ probably due to the combined influence of ~~the secondary oxidants from~~ ^{3}VL* and enhanced VL nitration in the presence of O_2~~, and nitrate photolysis products~~. ~~In our calculations, the increase in <OS_c> (except for VOCs and inorganic anions experiments; A9 to A12; Table S2) was lower than those in $^{\bullet}$OH- or triplet mediated oxidation of phenolics (e.g., phenol, guaiacol) measured using an aerosol mass spectrometer (Sun et al., 2010; Yu et al., 2014),~~Our measured <OS_c> for all experiments range from -0.28 to +0.12, while other studies on phenolic aqSOA formation reported <OS_c> ranging from -0.14 to +0.47 (Sun et al., 2010) and 0.04 to 0.74 (Yu et al., 2014). This is likely because we excluded contributions from ring-opening products, which may have higher OS_c values as these products

are not detectable in the positive ion mode. Thus, the <OS_c> in this study likely were lower estimates. In brief, the presence of ~~secondary oxidants from ^{3}VL* and~~ O_2 increased the normalized abundance of products and promoted the formation of more oxidized aqSOA. ~~These trends were reinforced in the presence of nitrate, indicating synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~Compared to N_2-saturated condition, the higher normalized abundance of nitrogen-containing products under air-saturated condition for VL+AN (at pH 4) suggests a potential enhancement of VL nitration in the presence of O_2.

Illumination of phenolic aromatic carbonyls with high molar absorptivities ($\varepsilon_{\lambda max}$) (~8 to 22 × 10^3 M^{-1} cm^{-1}) leads to an overall loss of light absorption but increased absorbance at longer wavelengths (> 350 nm), where the carbonyls did not initially absorb light (Smith et al., 2016). Fig. 3~~2~~a illustrates the changes in total absorbance from 350 to 550 nm of VL* and VL+AN under N_2- ~~air~~ and ~~N_2~~-air-saturated conditions. The absorption spectra of VL* under air- and N_2- saturated conditions (pH 4) at different time intervals are shown in Fig. S8. For both VL* and VL+AN, evident absorbance enhancement was observed under air-saturated conditions, while the absorbance changes under N_2-saturated conditions were minimal, consistent with the VL decay trends. This absorbance enhancement can be explained by the formation of oligomers with large, conjugated π-electron systems (Chang and Thompson, 2010) and hydroxylated products (Li et al., 2014; Zhao et al., 2015), in agreement with the observed reaction products. In this work, phenoxy radicals (in resonance with a carbon-centered cyclohexadienyl radical) can be generated from several processes such as the oxidation ~~(Vione et al., 2019)~~ of ground-state VL by ^{3}VL* via H-atom abstraction ~~(Huang et al., 2018)~~or electron transfer coupled with proton transfer from the phenoxyl radical cation or from solvent water (Neumann et al., 1986a, 1986b; Anastasio et al., 1996) and photoinduced O-H bond-breaking (Berto et al., 2016). ~~Moreover, ^{3}VL* can initiate H-atom abstraction from the CHO group of VL, generating ketyl radicals via Norrish type reactions (Vione et al., 2019).~~ Also, similar reactions can be initiated by $^{\bullet}$OH (Gelencsér et al., 2003; Hoffer et al., 2004; Chang and Thompson, 2010; Sun et al., 2010), which in this study can be generated from the reaction between ^{3}VL* and O_2, as well as nitrate photolysis. Trace amounts of H_2O_2 were likely formed during VL photodegradation (Li et al., 2014), similar to the case of other phenolic compounds (Anastasio et al., 1996). Oligomers can then form via the coupling of phenoxy radicals or phenoxy and cyclohexadienyl ~~ketyl~~ radicals (Sun et al., 2010; Yu et al., 2014~~Berto et al., 2016~~; Vione et al., 2019). Absorbance increase at > 350 nm has also been reported for photosensitized oxidation of phenol and 4-phenoxyphenol (De Laurentiis et al., 2013a, 2013b) and direct photolysis of tyrosine and 4-phenoxyphenol (Bianco et al., 2014) in which dimers have been identified as initial substrates. The continuous absorbance enhancement throughout 6 h of irradiation correlated with the observation of oligomers and nitrated compounds after irradiation. However, the increasing concentration of small organic acids (Fig. S5) throughout the experiments suggests that fragmentation, which results in the decomposition of initially formed oligomers and formation of smaller oxygenated products (Huang et al., 2018), is important at longer irradiation times. Overall, these trends establish that ~~secondary oxidants from ^{3}VL*and~~ O_2 ~~are~~ is necessary for the efficient formation of light-absorbing compounds from both VL* and VL+AN.

3.1.2 Effect of pH

The reactivity of $^3C^*$ (Smith et al., 2014, 2015, 2016), aromatic photonitration by nitrate (Machado and Boule, 1995; Dzengel et al., 1999; Vione et al., 2005; Minero et al., 2007), and N(III)-mediated VL photo-oxidation (Pang et al., 2019a) have been demonstrated to be pH-dependent. In this study, the effect of pH on VL photo-oxidation was investigated within the range of 2.5 to 4~~5~~, corresponding to typical cloud ~~(2-7)~~ pH values (2-7) (Pye et al., 2020). The decay rate~~s~~ constants for both VL* and VL+AN increased as pH decreased (VL* and VL+AN at pH 2.5: 1.6~~5~~ and 1.4~~3~~ times faster than at pH 4, respectively) (Fig. S3b). These differences in decay rate constants are small but statistically significant ($p < 0.05$). The pK_a for the VL triplet has been reported to be 4.0 (Smith et al., 2016). As there are a greater fraction of VL triplets that are protonated at pH 2.5 (0.96) than at pH 4 (0.5), it is possible that the pH dependence of the decay rate constants observed in this study is due to ^{3}VL* being more reactive in its protonated form. Smith et al. (2016) also observed a pH dependence for the direct photodegradation of VL (0.005 mM) (rate constants at pH ≤ 3 are ~two times lower than at pH ≥ 5) which they attributed to the sensitivity of the excimer of VL (i.e., the charge-transfer complex formed between an excited state VL molecule and a separate ground state VL molecule; Birks, 1973, Turro et al., 2010) to acid-base chemistry. The opposite trend observed in this study for 0.1 mM VL may be due to the reactivities of the protonated and neutral forms of the ^{3}VL* being dependent on the VL concentration (Smith et al., 2016). ~~For VL*, this pH trend indicates that ^{3}VL* are more reactive in their protonated form, which is opposite to that reported for 0.005 mM VL (Smith et al., 2016), likely due to the concentration dependence of the relative reactivities of protonated and neutral forms of ^{3}VL*.~~ It has been reported that the quantum yield for direct VL photodegradation is higher at pH 5 than at pH 2 for 0.005 mM VL, but they are not statistically different for 0.03 mM VL (Smith et al., 2016). ~~Also, increases in hydrogen ion concentration can enhance the formation of $HO_2^{\bullet}$ and H_2O_2 and in turn, $^{\bullet}OH$ formation (Du et al., 2011). In addition to these pH influences on VL* the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a) also contributed to the observed pH effects for VL+AN. At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}OH$ formation than that of NO_2^- in the near-UV region (Arakaki et al., 1999; Kim et al., 2014). The increased $^{\bullet}OH$ formation rates as pH decreases can lead to faster VL decay (Pang et al., 2019a). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}OH$ (Reactions 4 and 15; Table 1) (Pang et al., 2019a).~~ As pH decreases, the higher reactivity of ^{3}VL* and sensitivity of the excimer of VL to acid-base chemistry ~~HONO being the dominant N(III) species can lead~~ may have led to faster VL photo-oxidation. Similar to pH 4 experiments, comparable decay rate constants between VL* and VL+AN were also noted at pH < 4, again suggesting the predominant role of ^{3}VL* chemistry compared to nitrate, likely due to the high VL concentration (0.1 mM) used in this study.

As pH decreased, the normalized abundance of products, particularly oligomers and functionalized monomers, was higher for both VL* and VL+AN, further indicating that ^{3}VL* ~~are~~ may be more reactive in their protonated form. The most abundant products observed were a substituted VL ($C_{10}H_{10}O_5$; No. 4, Table S2) and VL dimer ($C_{16}H_{14}O_6$; No. 6~~5~~, ~~Table S3~~Table S2) at pH 4 and pH < 4, respectively (Figs. 1c-h). Furthermore, a tetramer ($C_{31}H_{24}O_{11}$) was observed only in VL* at pH 2.5. For VL+AN, the normalized abundance of nitrogen-containing compounds was also higher ~~increased~~ at lower pH

(Table ~~S~~2), likely due to increased $^{\bullet}OH$ and $^{\bullet}NO_2$ formation, which may be caused by the dependence of N(III) (NO_2^- + HONO) speciation on solution acidity (Pang et al., 2019a). At pH 3.3, half of N(III) exists as HONO (Fischer and Warneck, 1996; Pang et al., 2019a), which has a higher quantum yield for $^{\bullet}OH$ formation than that of NO_2^- in the near-UV region (Arakaki et al., 1999; Kim et al., 2014). Also, NO_2^-/HONO can generate $^{\bullet}NO_2$ via oxidation by $^{\bullet}OH$ (Reactions 4 and 10~~5~~; Table 1) (Pang et al., 2019a). At pH < 4, ^{3}VL* likely have higher reactivity as suggested by the increased normalized abundance of oligomers (e.g., $C_{16}H_{14}O_6$; No. 6, Table S2 and $C_{31}H_{24}O_{11}$) and N-containing compounds (e.g., $C_{16}H_{10}N_2O_9$; No. 3, Table S2 and $C_{13}H_{14}N_2O_{10}$) (Fig. 2). The potential imidazole derivative ($C_5H_5N_3O_2$; No. 5, Table S2) was observed only at pH 4, possibly due to the pH dependence of ammonium speciation (pK_a = 9.25). Imidazole formation requires the nucleophilic attack of ammonia on the carbonyl group (Yu et al., 2011), and at pH 4, the concentration of dissolved ammonia in VL+AN was about 10 or 30 times higher than that at pH 3 or pH 2.5, respectively. At different pH, the O:C and H:C ratios in VL* and VL+AN had no significant differences (Figs. S7c–d and S9), but molecules with higher O:C ratios (> 0.6) were more abundant at pH < 4. Accordingly, the $\langle OS_c \rangle$ at pH < 4 for both VL* and VL+AN were higher than that at pH 4, consistent with higher $\langle OS_c \rangle$ observed at pH 5 compared to pH 7 for the $^{\bullet}OH$-mediated photo-oxidation of syringol (Sun et al., 2010). Essentially, the higher reactivity of ^{3}VL* and predominance of HONO over nitrite at lower pH may have resulted in increased formation of products mainly composed of oligomers and functionalized monomers.

~~The h~~Higher absorbance enhancement for both VL* and VL+AN (Fig. 3~~2~~b) was observed as pH increased~~ may be attributed to redshifts and increased visible light absorption of reaction products (Pang et al., 2019a)~~. To determine whether the pH dependence is due to the acid-base chemistry of the products or of the reactions, we measured the pH dependence of the aqSOA formed from VL* at pH 4 and 2.5 over a range of pH conditions from 1.5 to 10.5 (Fig. S10). For both cases, the intensity of absorption at longer wavelengths significantly increased as the pH of the solutions was raised. Moreover, the comparable pH dependence of the two solutions suggests that the observed pH dependence may be attributed to the acid-base chemistry of the reactions, which may involve ^{3}VL* or the excimer of VL (Smith et al., 2016), as discussed earlier. ~~When a phenolic molecule deprotonates at higher pH, an ortho- or para- electron-withdrawing group, such as a nitro or aldehyde group, can attract a portion of the negative charge towards its oxygen atoms through induced and conjugated effects, leading to the extension of chromophore from the electron-donating group (e.g., $-O^-$) to the electron-withdrawing group via the aromatic ring (Carey, 2000; Williams and Fleming, 2008; Pang et al., 2019a). Hence, the delocalization of the negative charge in phenolates leads to significant redshifts (Mohr et al., 2013).~~

~~3.1.3 Effect of VOCs and inorganic anions~~

~~Aerosols are a complex mix of organic and inorganic compounds (Kanakidou et al., 2005). We explored the photo-oxidation behavior of VL, with and without nitrate, in the presence of VOCs (2-propanol, IPA) and inorganic anions (sodium bicarbonate, NaBC). For both VL* and VL+AN, there was no significant change in VL decay (Figs. S3c-d), and comparable absorbance enhancements (Figs. 2c-d) were observed upon the addition of IPA and NaBC. However, the characterization of reaction products revealed the distinct effects of these compounds on the photo-oxidation of VL. Both IPA and NaBC increased the~~

~~normalized abundance of products from VL* (by a factor of 2.4 and 1.4, respectively) and VL+AN (by a factor of ~4) (Table S2). The major product observed in VL*+IPA (Fig. S10a) was a dimer ($C_{16}H_{14}O_6$). Also, higher oligomers up to tetramers (e.g., $C_{31}H_{22}O_{12}$) not observed in VL* were noted. A possible explanation may be the additional generation of $^{\bullet}HO_2^-$ from the reaction of IPA with $^{\bullet}OH$ (Warneck and Wurzinger, 1988) (Reactions 16 and 17; Table 1), which can originate from ^{3}VL* or nitrate photolysis, inducing reactions such as oxidation and nitration. As discussed earlier, $^{\bullet}HO_2$ can form H_2O_2, a photolytic source of $^{\bullet}OH$ (Anastasio et al., 1996; Du et al., 2011). In the presence of IPA, the increase in normalized abundance of products (VL+AN+IPA: 3.8 times vs. VL*+IPA: 2.4 times; Table S2) and $<OS_c>$ (VL+AN+IPA: −0.13 to 0.08 vs. VL*+IPA: −0.16 to −0.10; Table S2) being more evident for VL+AN compared to VL* also supports the potential importance of reactions involving $^{\bullet}HO_2$ and nitrate photolysis products such as the secondary photochemistry between $^{\bullet}HO_2/O_2^{\bullet-}{}_{(aq)}$ and photoproduced $NO_{x(aq)}$ enhancing $HONO_{(g)}$ production from nitrate photolysis in the presence of dissolved aliphatic organic matter (Wang et al., 2021) as discussed in Sect. 3.1.1. This chemistry may have operated in VL+AN+IPA considering that $^{\bullet}HO_2/O_2^{\bullet-}$ may originate from multiple sources in this experiment: nitrate photolysis (Reaction 10; Table 1) (Jung et al., 2017; Wang et al., 2021), the reactions of ^{3}VL* in the presence of O_2 (see Sect. 3.1), or reaction of IPA with $^{\bullet}OH$ (Warneck and Wurzinger, 1988) (Reactions 16 and 17; Table 1). In other words, the role of nitrate in VL photo-oxidation is enhanced in the presence of IPA, likely due to additional $^{\bullet}HO_2/O_2^{\bullet-}$ formation. In VL+AN+IPA, nitrate photolysis likely converted $C_{16}H_{14}O_6$ (from VL*+IPA) to $C_{15}H_{12}O_8$ (Figs. S10a-b) via demethylation and then multiple hydroxylations. Nitrate photolysis generates $^{\bullet}OH$, and demethylation has been reported to be enhanced at high $^{\bullet}OH$ exposure (Gold et al., 1983). Moreover, alcohols can affect the structure of water, causing a localized patterning or organization that changes the solvation environment, which can account for reactivity enhancement in the presence of alcohol-containing solvents (Berke et al., 2019). Berke et al. (2019) has demonstrated that IPA and other alcohols (e.g., ethanol) can promote the production of light-absorbing compounds, i.e., imidazoles, from the reactions between glyoxal and ammonium sulfate. This phenomenon has been attributed to the formation of micro-heterogeneities of hydrated alcohol molecules in a complex solution environment composed of solvated sulfate ions and a mixture of reactants and products upon the addition of alcohols. As proposed by an earlier study (Onori and Santucci, 1996), if the water in the SOA-mimicking solutions exists in two forms, bulk and hydrating, the micro-heterogeneities may interact with water/nitrate matrix to sequester the reactants and products, concentrating them within a smaller effective solvent volume and consequently resulting in increased normalized abundance of products (Berke et al., 2019).~~

~~For NaBC which does not produce $^{\bullet}HO_2$ upon reactions with $^{\bullet}OH$ under air-saturated conditions (Gen et al., 2019b), the increased normalized abundance of products may be due to other reactions promoted by the carbonate radical ($CO_3^{\bullet-}$), which can be generated from the reactions of bicarbonate/carbonate with $^{\bullet}OH$ (Reactions 18 and 19; Table 1) (Neta et al., 1988; Wojnárovits et al., 2020) or ^{3}VL* (Reactions 20 and 21; Table 1) (Canonica et al., 2005). $CO_3^{\bullet-}$ is a selective oxidant that reacts with organic molecules at a lower rate than $^{\bullet}OH$ and readily reacts with electron-rich parts of phenols, aromatic amines, and sulfur-containing compounds (e.g., glutathione) through both electron transfer and H abstraction (Huang and Mabury, 2000; Wojnárovits et al., 2020). Similar to IPA, the enhancement of normalized abundance of products (VL+AN+NaBC: 4.3 times vs. VL*+NaBC: 1.4 times; Table S2) and $<OS_c>$ (VL+AN+NaBC: −0.13 to 0.08 vs. VL*+NaBC:~~

~~-0.16 to -0.11; Table S2) was more obvious for VL+AN+NaBC than VL*+NaBC, further underlining the contributions of nitrate photolysis products. For example, it has been reported that carbonate and bicarbonate can substantially increase the photodegradation of electron-rich compounds (e.g., catechol) by nitrate (Vione et al., 2009). Bicarbonate can enhance the photolysis of nitrate via a solvent-cage effect, reacting with photolysis-derived $^{\bullet}OH$ before it escapes the surrounding cage of the water molecules. This prevents the recombination of $^{\bullet}OH$ and $^{\bullet}NO_2$ inside the solvent cage that otherwise would yield back $NO_3^- + H^+$, which reduces the quantum yield of $^{\bullet}OH$ photoproduction (Bouillon and Miller, 2005). This scavenging of in-cage $^{\bullet}OH$ by bicarbonate would then hinder recombination, resulting in a higher generation rate of $CO_3^{\bullet-} + {}^{\bullet}OH$ with bicarbonate compared to $^{\bullet}OH$ alone without bicarbonate. However, in our experiments, NaBC did not cause any substantial change in the decay of VL for both VL* and VL+AN, although it promoted higher normalized abundance of products. The major product in VL*+NaBC was a functionalized monomer ($C_7H_4O_4$; No. 6, Table S3; Fig. S10c). Unlike VL*+IPA, no tetramers were observed in VL*+NaBC. Similar to VL+AN+IPA, the addition of NaBC to VL+AN resulted in trimers and a high-abundance dimer ($C_{15}H_{12}O_8$; No. 7, Table S3) (Figs. S10b and S10d). Overall, VL+AN+IPA had more oligomers while VL+AN+NaBC had more functionalized monomers (e.g., $C_8H_6O_4$; No. 8, Table S3). These findings suggest that aside from low pH (<4), the formation of oligomers from VL photo-oxidation can also be promoted by presence of VOCs and inorganic anions likely via the generation of radicals such as $^{\bullet}HO_2$ and $CO_3^{\bullet-}$ which can also interact with nitrate photolysis products.~~

~~The addition of IPA or NaBC to VL* resulted in products with higher O:C and H:C ratios (Figs. S11a and S11c). Although the products were more abundant in VL*+IPA than with NaBC, the distribution of their products in van Krevelen diagrams was rather similar. The increased in $\langle OS_c \rangle$ in the presence of IPA or NaBC was more significant for VL+AN than VL*, likely due to the interactions of nitrate photolysis products with $^{\bullet}HO_2$ and $CO_3^{\bullet-}$. For VL+AN, IPA and NaBC also increased the O:C and H:C ratios (Figs. S11b and S11d), and most products had OS_c >0, similar to less volatile and semi-volatile oxygenated organic aerosols (LV-OOA and SV-OOA) (Kroll et al., 2011).~~

3.1.3~~4~~ Distribution of potential BrC compounds

Figure S11~~2~~ plots the DBE values vs. number of carbons (n_c) (Lin et al., 2018) for the 50 most abundant products from pH 4 experiments under air-saturated conditions, along with reference to DBE values corresponding to fullerene-like hydrocarbons (Lobodin et al., 2012), cata-condensed polycyclic aromatic hydrocarbons (PAHs) (Siegmann and Sattler, 2000), and linear conjugated polyenes with a general formula C_xH_{x+2}. As light absorption by BrC requires uninterrupted conjugation across a significant part of the molecular structure, compounds with DBE/n_C ratios (shaded area in Fig. S11~~2~~) greater than that of linear conjugated polyenes are potential BrC compounds (Lin et al., 2018). Based on this criterion and the observed absorbance enhancement at > 350 nm (Fig. 3~~2~~), the majority of the 50 most abundant products from pH 4 experiments under air-saturated conditions were potential BrC chromophores composed of monomers and oligomers up to tetramers. However, as ESI-detected compounds in BB organic aerosols has been reported to be mainly molecules with n_c < 25 (Lin et al., 2018), there may be higher oligomers that were not detected in our reaction systems.

3.2 Effect of reactants concentration and molar ratios on the aqueous photo-oxidation of vanillin

To examine the influence of VL and nitrate concentration and their molar ratios on VL photo-oxidation, we also characterized the reaction products from lower [VL] (0.01 mM VL*; A104; Table S2), lower [VL] concentrations and an equal molar ratio of VL/nitrate (0.01 mM VL + 0.01 mM AN; A115; Table S2), and lower [VL] and 1:100 molar ratio of VL/nitrate (0.01 mM VL + 1 mM AN; A126; Table S2) at pH 4. The normalized abundance of products from low [VL] experiments (A104-A126; Table S2) were up to 1.4 times higher than that of high [VL] experiments (A5 and A7; Table S2). Nevertheless, the major products for both low and high [VL] experiments were functionalized monomers (Figs. 1c-d and S123a–c) such as $C_8H_6O_4$ (No. 7, Table S2) and $C_{10}H_{10}O_5$ (No. 4, Table S2). For both VL* and VL+AN, the contribution of < 200 m/z to the normalized abundance of products was higher at low [VL] than at high [VL], while the opposite was observed for > 300 m/z (Fig. S123d). This indicates that functionalization was favored at low [VL], as supported by the higher $\langle OS_c \rangle$, while oligomerization was the dominant pathway at high [VL], consistent with more oligomers or polymeric products reported from high phenols concentration (e.g., 0.1 to 3 mM) (Li et al., 2014; Slikboer et al., 2015; Ye et al., 2019). This is probably due to an increased concentration of phenoxy radicals (in resonance with a carbon-centered cyclohexadienyl radical) at high [VL], promoting radical-radical polymerization (Sun et al., 2010; Li et al., 2014). Moreover At low [VL], the contribution of < 200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, further suggesting the prevalence of functionalization for the former formation of more oxidized products. This may also be the reason why In addition, 1:1 VL/nitrate (A115; Table S2) had higher $\langle OS_c \rangle$ than 1:100 VL/nitrate (A126; Table S2) VL/nitrate, indicating the formation of more oxidized products, but had fewer N-containing compounds compared to the latter. A possible explanation is that at 1:1 VL/nitrate, VL may efficiently compete with NO_2^- for $^{\bullet}OH$ (from nitrate or nitrite photolysis, Reaction 4; Table 1) and indirectly reduce $^{\bullet}NO_2$. Similarly, hydroxylation has been suggested to be a more important pathway for 1:1 VL/nitrite than in 1:10 VL/nitrite (Pang et al., 2019a). This may also be the reason why 1:1 VL/nitrate (A15; Table S2) had higher $\langle OS_c \rangle$ than 1:100 (A16; Table S2) VL/nitrate but had fewer N-containing compounds compared to the latter. Moreover, the contribution of <200 m/z to the normalized abundance of products was higher for 1:1 than 1:100 VL/nitrate molar ratio, further suggesting the formation of more oxidized products. Fragmentation, which leads to the decomposition of previously formed oligomers and generation of small, oxygenated products such as organic acids (Huang et al., 2018), may also occur for the low [VL] experiments. However, its importance would likely be observed at longer irradiation times, similar to the high [VL] experiments.

3.3 Participation of ammonium in the aqueous photo-oxidation of vanillin

Imidazole and imidazole derivatives have been reported to be the major products of glyoxal and ammonium sulfate reactions at pH 4 (Galloway et al., 2009; Yu et al., 2011; Sedehi et al., 2013; Gen et al., 2018; Mabato et al., 2019). Here, we compared VL+AN and VL+SN at pH 4 in terms of reaction products and oxidative characteristics to confirm the participation of ammonium in the aqueous photo-oxidation of VL. In both experiments The normalized abundance of the products was

comparable i~~I~~n both experiments (A7 and A9~~13~~; Table ~~S~~2), with $C_{10}H_{10}O_5$ (No. 4, Table S2) as the most abundant product (Figs. 1d and S6a), but in VL+SN, there was a significant amount of a VL dimer ($C_{15}H_{12}O_8$; No. 8~~9~~, ~~Table S3~~Table S2). Moreover, the nitrogen-containing compounds were distinct. Aside from the potential imidazole derivative ($C_5H_5N_3O_2$; No. 5~~10~~, Table S2~~3~~), $C_8H_9NO_3$ (No. 2, Table S2) was also observed from VL+AN but only under N_2-saturated conditions (Fig. 1b), probably due to further oxidation by ~~secondary oxidants from~~ ^{3}VL*. The product analysis suggests the participation of ammonium in the aqueous-phase reactions. Ammonium salts are an important constituent of atmospheric aerosols particles (Jimenez et al., 2009), and reactions between dicarbonyls (e.g., glyoxal) and ammonia or primary amines have been demonstrated to form BrC (De Haan et al., 2009, 2011; Nozière et al., 2009; Shapiro et al., 2009; Lee et al., 2013; Powelson et al., 2014; Gen et al., 2018; Mabato et al., 2019). Relative to VL+AN, the products from VL+SN had higher O:C ratios (e.g., $C_7H_4N_2O_7$; No. 9~~11~~, Table S2~~3~~), OS_c, and $\langle OS_c \rangle$ values (Table ~~S~~2).

3.4 Oxidation of guaiacol by photosensitized reactions of vanillin and photolysis of nitrate

The oxidation of phenols by ^{3}C* has been mainly studied using non-phenolic aromatic carbonyls (Anastasio et al., 1996; Smith et al., 2014, 2015; Yu et al., 2014; Chen et al., 2020) and aromatic ketones (Canonica et al., 2000) as triplet precursors. Recently, ^{3}VL* have also been shown to oxidize syringol (Smith et al., 2016), a non-carbonyl phenol, although the reaction products remain unknown. In this section, we discussed the photo-oxidation of guaiacol (GUA), a lignocellulosic BB pollutant (Kroflič et al., 2015) that is also a non-carbonyl phenol, in the presence of VL (GUA+VL) or nitrate (GUA+AN). The dark experiments did not show any substantial loss of VL or GUA (Fig. S3ce). Due to its poor light absorption in the solar range, GUA is not an effective photosensitizer (Smith et al., 2014; Yu et al., 2014). Accordingly, the direct GUA photodegradation resulted in minimal decay, which plateaued after ~3 hours. However, in the presence of VL or nitrate, the GUA decay rate constant was ~~faster~~higher by 2.2 (GUA+VL) and 1.3~~2~~ (GUA+AN) times, respectively, than for direct GUA photodegradation. The enhancement of GUA decay rate constant in the presence of VL is statistically significant ($p < 0.05$), while that in the presence of AN is not ($p > 0.05$). This enhanced GUA decay rate constant may be due to the ~~following main reactions:~~ oxidation of GUA by ^{3}VL* ~~(or the secondary oxidants it generates upon reaction with O_2), oxidation by $^{\bullet}$OH produced from nitrate photolysis, or nitration by $^{\bullet}NO_2$ from nitrate photolysis~~. As mentioned earlier, ~~the~~ ^{3}VL* chemistry appears to be more important than that of nitrate photolysis even at 1:10 molar ratio of VL/nitrate on account of the much higher molar absorptivity of VL compared to that of nitrate (Fig. S1) and the high VL concentration (0.1 mM) used in this study. However, the apparent quantum efficiency of GUA photodegradation (φ_{GUA}) in the presence of nitrate ($1.3 \times 10^{-2} \pm 2.9 \times 10^{-3}$) is ~14 times larger than that in the presence of VL ($9.0 \times 10^{-4} \pm 4.0 \times 10^{-4}$), suggesting that nitrate-mediated photo-oxidation of GUA is more efficient than that by photosensitized reactions of VL (see Text S7 for more details). The decay of VL in GUA+VL (A14~~8~~; Table ~~S~~2) was 3 times slower than that of VL* (A5; Table ~~S~~2), which may be due to competition between ground-state VL and GUA for reactions with ^{3}VL* ~~(or the secondary oxidants it generates upon reaction with O_2)~~ or increased conversion of ^{3}VL* back to the ground state through the oxidation of GUA (Anastasio et al., 1996; Smith et al., 2014). The corresponding absorbance changes for the GUA experiments (Fig. 3~~1~~ce) were consistent with the observed decay trends. The minimal absorbance changes

for the direct GUA photodegradation also plateaued after ~3 hours. Moreover, the difference between GUA photo-oxidation in the presence of VL or nitrate was more evident, with the former showing much higher absorbance enhancement. ~~Similarly,~~ Yang et al. (2021) also observed greater light absorption during nitrate-mediated photo-oxidation relative to direct GUA photodegradation.

For the direct GUA photodegradation, GUA+VL, and GUA+AN, the normalized abundance of products was calculated only for GUA+VL (2.2; Table ~~S~~2), as the GUA signal from the UHPLC-qToF-MS in the positive ion mode was weak, which may introduce large uncertainties during normalization. Nonetheless, the number of products detected from these experiments (178, 266, and 844 for the direct GUA photodegradation, GUA+AN, and GUA+VL, respectively) corroborates the kinetics and absorbance results. The major products (Fig. 4~~3~~a) from the direct photodegradation of GUA were $C_{14}H_{14}O_4$ (No. 10~~9~~, ~~Table S3~~Table S2), a typical GUA dimer, and a trimer ($C_{21}H_{20}O_6$; No. 11~~20~~, ~~Table S3~~Table S2) which likely originated from photoinduced O-H bond-breaking (Berto et al., 2016). In general, higher absolute signals ~~areas~~ were~~as~~ noted for oligomers (e.g., $C_{14}H_{14}O_4$, No. 10 and $C_{21}H_{20}O_6$, No. 11, Table S2) and hydroxylated products (e.g., $C_7H_8O_4$) in both GUA+VL and GUA+AN, similar to those observed from GUA oxidation by triplets of 3,4-dimethoxybenzaldehyde (DMB; a non-phenolic aromatic carbonyl) or $^{\bullet}OH$ (from H_2O_2 photolysis) (Yu et al., 2014). In contrast to the GUA aqSOA reported by Yu et al. (2014), the photo-oxidation of GUA in this study yielded nitrated compounds (e.g., $C_9H_{14}N_2O_6$, $C_{11}H_{14}N_2O_9$) from GUA+AN and VL dimers (e.g., $C_{16}H_{12}O_6$) from GUA+VL. However, based on a recent work on the aqueous photo-oxidation of guaiacyl acetone (another aromatic phenolic carbonyl) by DMB triplets, the hydroxylation and dimerization of DMB can also contribute to aqSOA (Jiang et al., 2021). The contributions from DMB-participated reactions were only minor due to the low initial DMB concentration (0.005 mM). Relative to GUA+AN, higher signals for dimers such as $C_{14}H_{14}O_4$ (No. 10, Table S2) and $C_{16}H_{12}O_6$ were noted in GUA+VL, possibly due to both GUA and ground-state VL being available as oxidizable substrates for ^{3}VL* ~~and the secondary oxidants it can generate~~. Also, a potential GUA tetramer ($C_{28}H_{24}O_8$, No. 12~~21~~, Table S2~~3~~) was observed only in GUA+VL, consistent with higher oligomer formation from the triplets-mediated photo-oxidation of phenolics relative to $^{\bullet}OH$-assisted photo-oxidation (Yu et al., 2014). In general, the products from the direct GUA photodegradation, GUA+VL, and GUA+AN had similar OS_c values (-0.5 to 0.5) (Figs. 4~~3~~b–d), falling into the criterion of BBOA and SV-OOA (Kroll et al., 2011). In this work, ~~efficient~~ GUA photo-oxidation was observed in the presence of VL or~~and~~ AN, forming aqSOA composed of oligomers, hydroxylated products, and nitrated compounds (for GUA+AN). The higher product signals from GUA+VL compared to GUA+AN is likely due to the availability of both GUA and ground-state VL as aqSOA precursors.

~~**3.5 Photo-oxidation pathways of vanillin via direct photosensitization and in the presence of nitrate**~~

~~The most probable pathways of direct photosensitized and nitrate-mediated photo-oxidation of VL were proposed (Fig. 4). In Scheme 1 (pH 4 and pH <4 under air-saturated conditions), ^{3}VL* and $^{\bullet}OH$ (from ^{3}VL* or nitrate photolysis) can initiate H-atom abstraction to generate phenoxy or ketyl radicals (Huang et al., 2018; Vione et al., 2019). At pH 4, ring-opening products~~

~~(Fig. S5) from fragmentation in both VL* and VL+AN may have reacted with VL or dissolved ammonia to generate $C_{10}H_{10}O_5$ (Pang et al., 2019b) and a potential imidazole derivative ($C_5H_5N_3O_2$), respectively. Moreover, nitrate photolysis products promoted functionalization and nitration (e.g., $C_{16}H_{10}N_2O_9$). At pH <4, the reactivity of ^{3}VL* increased as suggested by the abundance of oligomers (e.g., $C_{16}H_{14}O_6$) and increased normalized abundance of N-containing compound~~

~~In Scheme 2 (pH 4, IPA or NaBC, under air-saturated conditions), additional radicals generated ($^{\bullet}HO_2$ and $CO_3^{\bullet-}$) likely promoted more reactions. An abundant dimer ($C_{16}H_{14}O_6$) and higher oligomers (e.g., tetramers, $C_{31}H_{22}O_{12}$) were identified in VL*+IPA, possibly due to $^{\bullet}HO_2$ initiated reactions, while a functionalized monomer ($C_7H_4O_4$) was abundant in VL*+NaBC. In general, nitrate enhanced both oligomerization and functionalization in VL+IPA or VL+NaBC. In VL+AN+IPA, $C_{15}H_{12}O_8$ likely originated from $C_{16}H_{14}O_6$ via demethylation and multiple hydroxylations. In VL+AN+NaBC, $C_8H_6O_4$ was possibly generated via H atom abstraction from $^{\bullet}OCH_3$ by $^{\bullet}OH$, and further addition with O_2 is energy barrierless (Priya and Lakshmipathi, 2017; Sun et al., 2019), generating a hydroperoxide ($^{\bullet}OCH_2OOH$) that readily decompose to form $^{\bullet}OCH_2O^{\bullet}$ and $^{\bullet}OH$ (Yaremenko et al., 2016). $^{\bullet}OCH2O^{\bullet}$ is finally transformed into $^{\bullet}OCHO$ with the elimination of HO_2 in the presence of O_2 (Sun et al., 2019). Moreover, the abundance of $C_{15}H_{12}O_8$ was higher in VL+AN+NaBC than in VL*+NaBC.~~

4 Conclusions and atmospheric implications

This study shows that the photo-oxidation of VL via its direct photosensitized reactions and in the presence of nitrate can generate aqSOA composed of oligomers, functionalized monomers, oxygenated ring-opening products, and nitrated compounds (for nitrate-mediated reactions). The characterization of products presented in this work complements earlier studies (e.g., Smith et al. 2014, 2015, 2016) that mainly discussed the kinetics and aqSOA yield of triplet-driven oxidation of phenols. Although nitrate did not substantially affect the VL decay rate~~s~~ constants, likely due to much higher molar absorptivity of VL than nitrate and high VL concentration used in this work, the presence of nitrate promoted functionalization and nitration, indicating the significance of nitrate photolysis in this aqSOA formation pathway. ~~While~~This work demonstrates that -nitration, which is ~~can be~~ an important process for producing light-absorbing organics or BrC (Jacobson, 1999; Kahnt et al., 2013; Mohr et al., 2013; Laskin et al., 2015; Teich et al., 2017; Li et al., 2020), ~~its effect on~~can also affect the aqueous-phase processing of triplet-generating aromatics ~~has not yet been examined in detail~~. On a related note, a recent work (Ma et al., 2021) mimicking phenol oxidation by DMB (a non-phenolic aromatic carbonyl) triplets in more concentrated conditions in aerosol liquid water (ALW) showed that significantly higher AN concentration (0.5 M) increased the photodegradation rate constant for guaiacyl acetone (an aromatic phenolic carbonyl with high Henry's law constant, 1.2×10^6 M atm^{-1}; McFall et al., 2020) by > 20 times which was ascribed to $^{\bullet}OH$ formation from nitrate photolysis (Brezonik and Fulkerson-Brekken, 1998; Chu and Anastasio, 2003). The same study also estimated that reactions of phenols with high Henry's law constants (10^6 to 10^9 M atm^{-1}) can be important for SOA formation in ALW, with mechanisms mainly governed by $^3C^*$ and $^1O_2$ (Ma et al., 2021). Likewise, Zhou et al. (2019) reported that the direct photodegradation of acetosyringone was faster by about 6 times in the presence of 2 M $NaClO_4$. However, the opposite was noted for the photodegradation of VL in sodium sulfate or sodium nitrate, which would

occur slower (~2 times slower in 0.5 M sodium sulfate and ~10 times slower in 0.124 M sodium nitrate) in ALW relative to dilute aqueous phase in clouds. These suggest that the nature of inorganic ions may have an essential role in the photodegradation of organic compounds in the aqueous phase (Loisel et al., 2021).

Furthermore, a potential imidazole derivative observed from the VL+AN (A7; Table ~~S~~2) experiment suggests that ammonium may participate in aqSOA formation from the photo-oxidation of phenolic aromatic carbonyls. Also, the oligomers from these reaction systems may be rather recalcitrant to fragmentation based on their high abundance, even at the longest irradiation time used in this study. Nonetheless, the increasing concentration of small organic acids over time implies that fragmentation becomes important at extended irradiation times. Aromatic carbonyls and nitrophenols have been reported to be the most important classes of BrC in cloud water heavily affected by biomass burning in the North China Plain (Desyaterik et al., 2013). Correspondingly, the most abundant products from our reaction systems (pH 4, air-saturated solutions) are mainly potential BrC chromophores. These suggest that aqSOA generated in cloud/fog water from the oxidation of biomass burning aerosols via direct photosensitized reactions and nitrate photolysis products can impact aerosol optical properties and radiative forcing, particularly for areas where biomass burning is intensive.

Our results indicate that the photo-oxidation of VL is influenced by O_2~~secondary oxidants from VL triplets~~, pH,~~ the presence of VOCs and inorganic anions,~~ and reactants concentration and molar ratios. ~~Compared to~~ Under N_2-saturated conditions,~~–~~ the absence of O_2 likely hindered the secondary steps in VL decay (e.g., reaction of ketyl radical and O_2), regenerating VL as suggested by the minimal VL decay~~more efficient VL photo-oxidation was observed under air-saturated conditions (O_2 is present), which can be attributed to the generation of secondary oxidants (e.g., $^1O_2$, $O_2^{\bullet-}$/$^{\bullet}HO_2$, $^{\bullet}OH$) from $^3VL^*$~~. ~~Further enhancement of VL photo-oxidation under air-saturated conditions in the presence of nitrate indicates synergistic effects between secondary oxidants from VL triplets and nitrate photolysis products.~~In contrast, $^3VL^*$-initiated reactions proceeded rapidly under air-saturated conditions (O_2 is present) as indicated by higher VL decay rate constant and increased normalized abundance of products. For pH 4 experiments, the presence of both O_2 and nitrate resulted in the highest normalized abundance of products (including N-containing compounds) and <OS_c>, which may be due to O_2 promoting VL nitration. Nevertheless, further work is necessary to assess the effect of O_2 on the reactive intermediates involved in $^3VL^*$-driven photo-oxidation and elucidate the mechanisms of VL photo-oxidation under air-saturated conditions. ~~-~~Additionally, the formation of oligomers from VL photo-oxidation was ~~observed to be~~ promoted at low pH (< 4)~~ or in the presence of IPA/NaBC, which likely generated additional radicals such as $^{\bullet}HO_2$ and $CO_3^{\bullet-}$~~. ~~As aerosols comprise more complex mixtures of organic and inorganic compounds, it is worthwhile to explore the impacts of other potential aerosol constituents on aqSOA formation and photo-oxidation studies. This can also be beneficial in understanding the interplay among different reaction mechanisms during photo-oxidation.~~Moreover, l~~L~~ow VL concentration favored functionalization, while oligomerization prevailed at high VL concentration, consistent with past works (Li et al., 2014; Slikboer et al., 2015; Ye et al., 2019). Hydroxylation was observed to be important for equal molar ratios of VL and nitrate, likely due to VL competing with nitrite for $^{\bullet}OH$. The oxidation of guaiacol, a non-carbonyl phenol, by photosensitized reactions of vanillin was also shown to be ~~more~~less~~-~~ efficient than that by nitrate photolysis products based on its lower apparent quantum efficiency.

In this study, we investigated reactions of VL and nitrate at concentrations in cloud/fog water. The concentrations of VL and nitrate can be significantly higher in aqueous aerosol particles. As a major component of aerosols, the concentration of nitrate can be as high as sulfate (Huang et al., 2014). More studies should then explore the direct photosensitized oxidation and nitrate-mediated photo-oxidation of other biomass burning-derived phenolic aromatic carbonyls, particularly those with high molar absorption coefficients and can generate $^{3}C^{*}$. The influences of reaction conditions should also be investigated to better understand the oxidation pathways. As aerosols comprise more complex mixtures of organic and inorganic compounds, it is worthwhile to explore the impacts of other potential aerosol constituents on aqSOA formation and photo-oxidation studies. This can also be beneficial in understanding the interplay among different reaction mechanisms during photo-oxidation. Considering that biomass burning emissions are expected to increase continuously, further studies on these aqSOA formation pathways are strongly suggested.

Data availability.

The data used in this publication are available to the community and can be accessed by request to the corresponding author.

Author contributions.

BRGM designed and conducted the experiments; YL provided assistance in measurements and helped to analyze experimental data; YJ provided assistance in measurements; BRGM, YL, and CKC wrote the paper. All co-authors contributed to the discussion of the manuscript.

Competing interests.

The authors declare that they have no conflict of interest.

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Table 1. List of reactions involving reactive species relevant to this study.

No.	Reactions	References
1	$NO_3^- + h\nu \rightarrow {}^{\bullet}NO_2 + O^-$; $\phi = 0.01$	Vione et al., 2006; ~~Scharko et al., 2014~~Benedict et al., 2017
2	$O^- + H_3O^+ \leftrightarrow {}^{\bullet}OH + H_2O$	
3	$NO_3^- + h\nu \rightarrow NO_2^- + O(^3P)$; $\phi = 0.\cancel{0}011$	
4	$NO_2^- + {}^{\bullet}OH \rightarrow {}^{\bullet}NO_2 + OH^-$ ($k = 1.0 \times 10^{10}$ M^{-1} s^{-1})	Mack and Bolton, 1999; Pang et al., 2019a
5	$O_2^{\bullet-} + NO_2^- + 2H^+ \rightarrow {}^{\bullet}NO_2 + H_2O_2$	Vione et al., 2001; Pang et al., 2019a
6	$NO_2^- + h\nu \rightarrow {}^{\bullet}NO + O^-$; $\phi_{OH,300} = 6.7\ (\pm\ 0.9)\%$	Fischer and Warneck, 1996; Mack and Bolton, 1999; Pang et al., 2019a
7	${}^{\bullet}NO + O_2 \leftrightarrow {}^{\bullet}ONOO$	Goldstein and Czapski, 1995~~a~~; Pang et al., 2019a
8	${}^{\bullet}ONOO + {}^{\bullet}NO \rightarrow ONOONO$	
9	$ONOONO \rightarrow 2{}^{\bullet}NO_2$	
~~10~~	~~$NO_3^- + h\nu \rightarrow {}^{\bullet}NO_2 + OH$ (reactions 1 & 2) $\rightarrow HOONO \xrightleftharpoons{h\nu} {}^{\bullet}NO + {}^{\bullet}HO_2$ ($pK_a = 6.8$)~~	~~Goldstein et al., 2005; Vione et al., 2005; Sturzbecher-Höhne et al., 2009; Abida et al., 2011; Wang et al., 2021~~
~~11~~	~~${}^{\bullet}HO_2 \rightleftharpoons H^+ + O_2^{\bullet-} \xrightleftharpoons{NO_2} OONO_2^- \xrightarrow{H_2O} O_2 + NO_2^- \xrightleftharpoons{+H^+} HONO$ ($pK_a = 4.8$) ($pK_a = 3.2$)~~	~~mmel et al., 1990; Goldstein et al., 1998; Wang et al., 21~~
~~12~~	~~$NO_2^- + HOONO_2 \rightleftharpoons OONO_2^- + H^+$ $O_2^{\bullet-} + NO \rightleftharpoons OONO^-$ ($pK_a = 5.9$) $\xrightarrow{HOONO} NO_3^-$~~	~~Goldstein and Czapski, 1995b; Wang et al., 2021~~
~~13~~	~~$HNO_2 + h\nu \rightarrow {}^{\bullet}NO + OH$; $\phi_{OH,300} = 36.2(\pm 4.7)\%$~~	~~Fischer and Warneck, 1996; Kim et al., 2014; Pang et al., 2019a~~
~~14~~	~~$HOONO \rightarrow {}^{\bullet}NO_2 + {}^{\bullet}OH$ ($k = 0.35 \pm 0.03$ s^{-1})~~	~~Goldstein et al., 2005; Pang et al., 2019a~~
1~~5~~	$HNO_2 + {}^{\bullet}OH \rightarrow {}^{\bullet}NO_2 + H_2O$ ($k = 2.6 \times 10^9$ M^{-1} s^{-1})	Kim et al., 2014; Pang et al., 2019a
~~16~~	~~$(CH_3)_2CHOH^- + {}^{\bullet}OH \rightarrow (CH_3)_2COH^{\bullet-} + H_2O$~~	~~Warneck and Wurzinger, 1988; Pang et al., 2019a~~
~~17~~	~~$(CH_3)_2COH^{\bullet-} + O_2 \rightarrow (CH_3)_2CO^- + {}^{\bullet}HO_2$~~	
~~18~~	~~${}^{\bullet}OH + HCO_3^- \rightarrow CO_3^{\bullet-} + H_2O$ ($k = 8.5 \times 10^6$ M^{-1} s^{-1})~~	~~Wojnárovits et al., 2020~~
~~19~~	~~${}^{\bullet}OH + CO_3^{2-} \rightarrow CO_3^{\bullet-} + OH^-$ ($k = 3.9 \times 10^8$ M^{-1} s^{-1})~~	
~~20~~	~~${}^3C^* + HCO_3^- \rightarrow CO_3^{\bullet-} + H^+ + C^{\bullet-}$ ($k = 10^6$–10^7 M^{-1} s^{-1}; ${}^3C^*$: triplet aromatic ketones)~~	~~Canonica et al., 2005~~
~~21~~	~~${}^3C^* + CO_3^{2-} \rightarrow CO_3^{\bullet-} + C^{\bullet-}$ ($k = 10^6$–10^7 M^{-1} s^{-1}; ${}^3C^*$: triplet aromatic ketones)~~	

Table 2. Reaction conditions, initial VL (and GUA) decay rate constants, normalized abundance of products, and average carbon oxidation state ($<OS_c>$) in each experiment. Except where noted, the reaction systems consisted of VL (0.1 mM); GUA (0.1 mM), AN (1 mM); sodium nitrate (SN) (1 mM) under air-saturated conditions after 6 h of simulated sunlight irradiation. Analyses were performed using UHPLC-qToF-MS equipped with an ESI source and operated in the positive ion mode.

Exp no.	pH	Reaction conditions	Initial VL (and GUA) decay rate constants (min^{-1})[b]	Ratio of 50 most abundant products to total products[c]	Normalized abundance of products[d]	Normalized abundance of N-containing compounds[d]	$<OS_c>$[e] (OS_c of VL: -0.25)
A1	2.5	VL*	$2.0 \times 10^{-2} \pm 5.8 \times 10^{-5}$	0.59	1.7 ± 0.16	N/A	-0.05
A2		VL+AN	$1.7 \times 10^{-2} \pm 7.3 \times 10^{-4}$	0.63	1.4 ± 0.19	5.3×10^{-2}	-0.04
A3	3	VL*	$1.5 \times 10^{-2} \pm 4.2 \times 10^{-4}$	0.53	1.9 ± 0.33	N/A	-0.04
A4		VL+AN	$1.5 \times 10^{-2} \pm 2.3 \times 10^{-4}$	0.56	1.9 ± 0.30	3.6×10^{-2}	-0.05
A5	4	VL*	$1.2 \times 10^{-2} \pm 5.9 \times 10^{-4}$	0.58	0.26 ± 0.42	N/A	-0.16
A6		VL* (N_2-saturated)	$3.2 \times 10^{-3} \pm 1.1 \times 10^{-3}$	0.96	4.7×10^{-2} ± 0.0027	N/A	-0.24
A7		VL+AN	$1.2 \times 10^{-2} \pm 8.8 \times 10^{-4}$	0.53	0.37 ± 0.38	1.7×10^{-2}	-0.13
A8		VL+AN (N_2-saturated)	$1.9 \times 10^{-3} \pm 9.2 \times 10^{-5}$	0.89	0.12 ± 0.0095	6.3×10^{-3}	-0.21
A9		VL+SN	N/A	0.51	0.42 ± 0.33	1.7×10^{-2}	-0.07
A10		VL* (0.01 mM)[a]	N/A	0.90	0.37 ± 0.018	N/A	-0.07
A11		VL (0.01 mM) + AN (0.01 mM)	N/A	0.77	0.40 ± 0.074	8.6×10^{-3}	0.12
A12		VL (0.01 mM) + AN	N/A	0.42	0.45 ± 0.025	1.2×10^{-2}	-0.06
A13		GUA only	$6.2 \times 10^{-3} \pm 2.5 \times 10^{-4}$	0.77	N/A	N/A	-0.28
A14		GUA+VL	GUA: $1.4 \times 10^{-2} \pm 4.0 \times 10^{-4}$ VL: $4.3 \times 10^{-3} \pm 2.2 \times 10^{-4}$	0.60	2.2 ± 0.47	N/A	-0.27
A15		GUA+AN	$8.0 \times 10^{-3} \pm 2.9 \times 10^{-3}$	0.77	N/A	N/A	-0.26

[a]Irradiation time for VL* (0.01 mM, A10) was 3 h. [b]The data fitting was performed in the initial linear region. Each value is the average of results from triplicate experiments. Errors represent one standard deviation. Kinetic measurements were not performed for experiments marked with N/A. [c]Ratio of the normalized abundance of the 50 most abundant products to that of total products, except for direct GUA photodegradation, GUA+VL, and GUA+AN (A13–15) whose ratios are based on the absolute signals of products. [d]The normalized abundance of products was calculated using Eq. 2. The samples for experiments without nitrate (marked with N/A) were not analyzed for N-containing compounds. For the GUA experiments, the normalized abundance of products was calculated only for GUA+VL as the GUA signal from the UHPLC-qToF-MS in the positive ion mode was weak, which may introduce large uncertainties during normalization. [e]$<OS_c>$ of the 50 most abundant products.

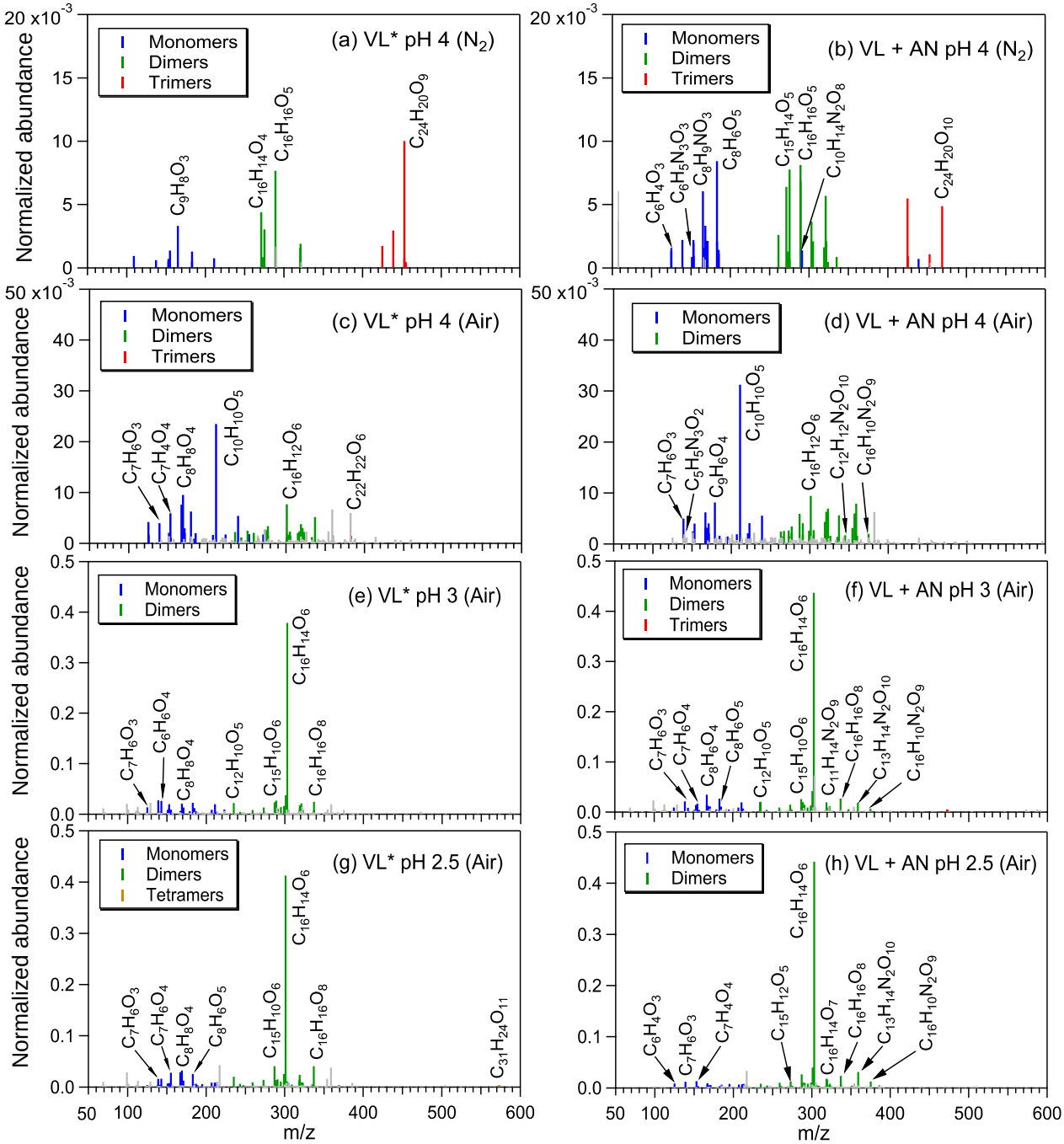

Figure 1. Reconstructed mass spectra of assigned peaks from (a) VL* pH 4 (N_2-saturated; A6), (b) VL+AN pH 4 (N_2-saturated; A8), (c) VL* pH 4 (air-saturated; A5), (d) VL+AN pH 4 (air-saturated; A7), (e) VL* pH 3 (air-saturated; A3), (f) VL+AN pH 3 (air-saturated; A4), (g) VL* pH 2.5 (air-saturated; A1), and (h) VL+AN pH 2.5 (air-saturated; A2) after 6 h of simulated

sunlight irradiation. The normalized abundance of products was calculated from the ratio of the peak area of the product to that of VL (Eq. 2). The 50 most abundant products contributed more than half of the total normalized abundance of products, and they were identified as monomers (blue), dimers (green), trimers (red), and tetramers (orange). Grey peaks denote peaks with low abundance or unassigned formula. Examples of high-intensity peaks were labeled with the corresponding neutral formulas. Note the different scales on the y-axes.

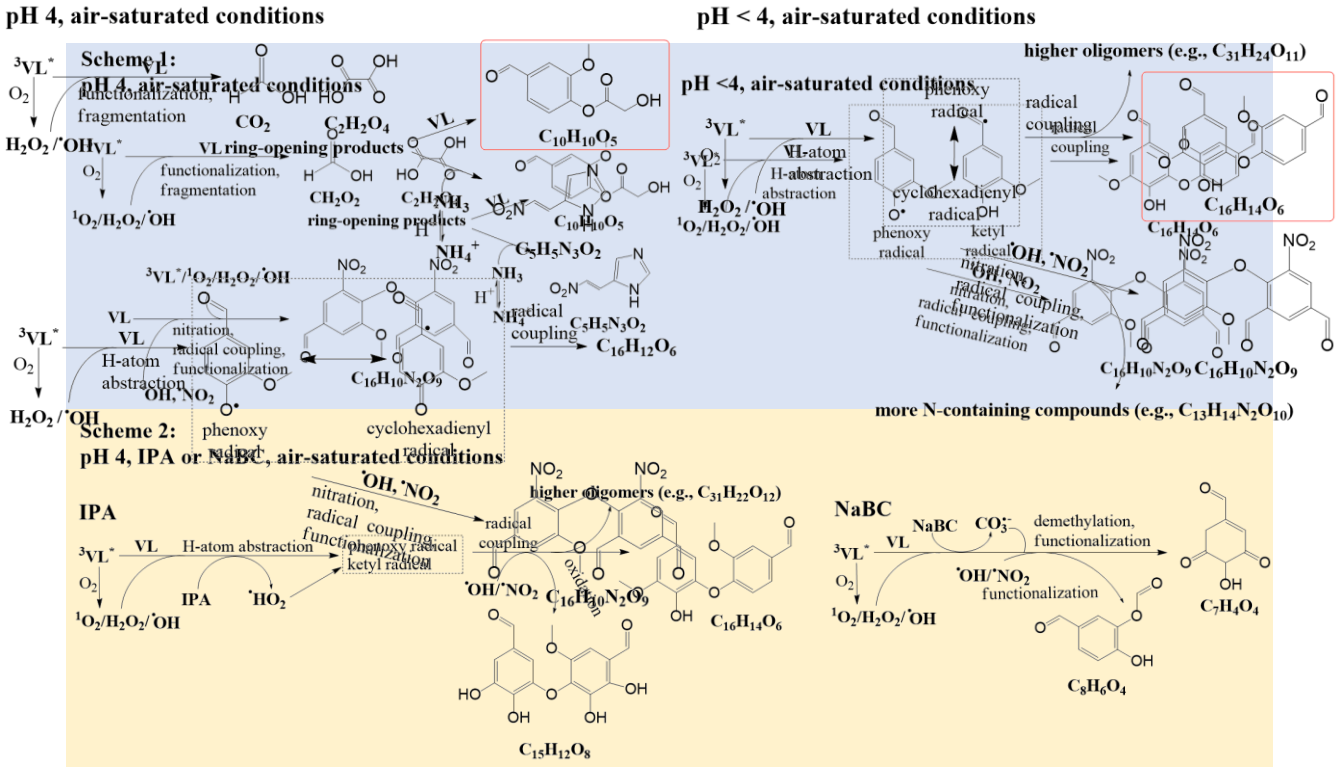

Figure 2. Potential aqSOA formation pathways via direct photosensitized oxidation of VL and nitrate-mediated VL photo-oxidation at pH 4 and pH < 4 under air-saturated conditions. Product structures were proposed based on the molecular formulas, DBE values, and MS/MS fragmentation patterns. The molecular formulas presented were the major products detected using UHPLC-qToF-MS in positive ESI mode. The highlighted structures are the most abundant product for each condition.

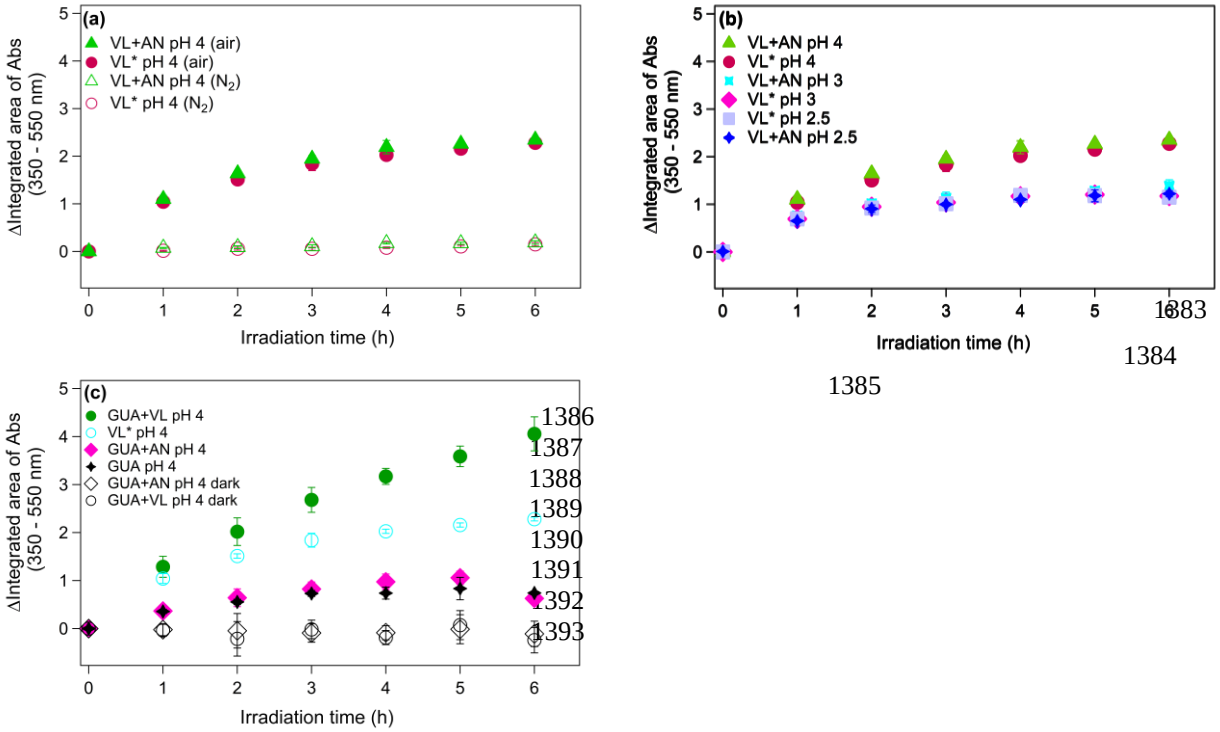
(a)
VL+AN pH 4 (air)
VL* pH 4 (air)
VL+AN pH 4 (N2)
VL* pH 4 (N2)
ΔIntegrated area of Abs (350 - 550 nm)
Irradiation time (h)
(b)
VL+AN pH 4
VL* pH 4
VL+AN pH 3
VL* pH 3
VL* pH 2.5
VL+AN pH 2.5
ΔIntegrated area of Abs (350 - 550 nm)
Irradiation time (h)
(c)
GUA+VL pH 4
VL* pH 4
GUA+AN pH 4
GUA pH 4
GUA+AN pH 4 dark
GUA+VL pH 4 dark
ΔIntegrated area of Abs (350 - 550 nm)
Irradiation time (h)

Figure 3~~2~~. (a~~-~~c~~d~~) Increase in light absorption under different experimental conditions for direct photosensitized oxidation of VL (VL*) and nitrate-mediated VL photo-oxidation (VL+AN): (a) ~~Effect of secondary oxidants from VL triplets on~~ VL* and VL+AN at pH 4 under N_2- (A6, A8) and air-saturated (A5, A7) conditions. (b) Effect of pH on VL* and VL+AN at pH 2.5 (A1, A2), 3 (A3, A4), and 4 (A5, A7) under air-saturated conditions. ~~(c) Effect of VOCs and inorganic anions: IPA (A9) and NaBC (A10) on VL* at pH 4 under air-saturated conditions. (d) Effect of VOCs and inorganic anions: IPA (A11) and NaBC (A12) on VL+AN at pH 4 under air-saturated conditions.~~ (c~~e~~) Increase in light absorption during direct GUA photodegradation (A13~~7~~) and photo-oxidation of GUA in the presence of VL (GUA+VL; A14~~8~~) or nitrate (GUA+AN; A15~~9~~) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. Error bars represent one~~1~~ standard deviation; most error bars are smaller than the markers.

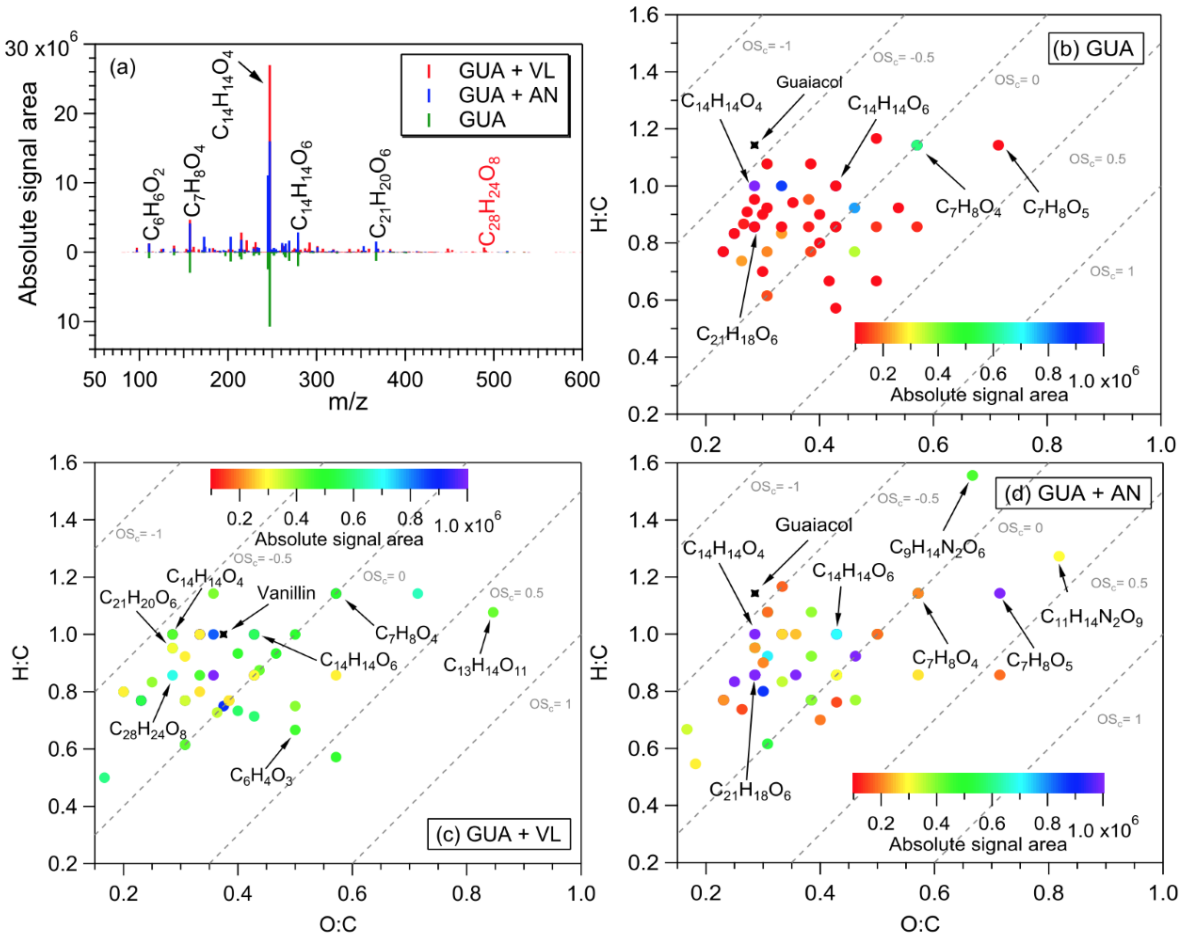

Figure 4~~3~~. (a) Reconstructed mass spectra of assigned peaks from the direct GUA photodegradation (A13~~7~~) and photo-oxidation of GUA in the presence of VL (GUA+VL; A14~~8~~) or nitrate (GUA+AN; A15~~9~~) at pH 4 under air-saturated

conditions after 6 h of simulated sunlight irradiation. The y-axis is the absolute signal area of the products. Examples of high-intensity peaks were labeled with the corresponding neutral formulas. (b-d) van Krevelen diagrams of the 50 most abundant products from the (b) direct photodegradation of GUA (A13~~7~~), (c) GUA+VL (A14~~8~~), and (d) GUA+AN (A15~~9~~) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. The color bar denotes the absolute signal area. The grey dashed lines indicate the carbon oxidation state values (e.g., OS_c =-1, 0, and 1).

~~**Figure 4.** Potential photo-oxidation pathways of VL via direct photosensitized reactions and in the presence of nitrate to illustrate the effects of secondary oxidants from VL triplets, pH, and the presence of VOCs (IPA) and inorganic anions (NaBC). Product structures were proposed based on the molecular formulas, DBE values, and MS/MS fragmentation patterns. The molecular formulas presented were the most abundant products or products with a significant increase in normalized abun~~

Supplementary material

Aqueous SOA formation from the photo-oxidation of vanillin: Direct photosensitized reactions and nitrate-mediated reactions

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Text S1. Materials.

Initial solutions of 0.1 mM vanillin (VL, Acros Organics, 99%, pure), 0.1 mM guaiacol (GUA, Sigma Aldrich, ≥ 98.0%), 1 mM ammonium nitrate (AN, Acros Organics, 99+%, for analysis), and 1 mM sodium nitrate (SN, Sigma-Aldrich, ≥ 99.5%)~~, 1 mM of 2-propanol (IPA, Optima LC/MS grade), and 1 mM of sodium bicarbonate (NaBC, Fisher BioReagents, 99.7-100.3%)~~ were prepared in Milli-Q water. The pH values of the samples were adjusted using sulfuric acid (H_2SO_4; Acros Organics, ACS reagent, 95% solution in water).

Text S2. UV-Vis spectrophotometric analyses.

The absorbance changes for all samples were characterized using a UV-Vis spectrophotometer (UV-3600, Shimadzu Corp., Japan). The absorbance values from 200 to 700 nm were recorded instantly after sample collection, and measurements were done in triplicate. Absorbance enhancements were calculated as the change in the integrated area of absorbance from 350 to 550 nm. The increase of light absorption at this wavelength range, where VL and GUA did not initially absorb light, suggests the formation of light-absorbing compounds (Zhou et al., 2019).

Text S3. UHPLC-PDA analyses.

An ultra-high performance liquid chromatography system (UHPLC, Waters Acquity H-Class, Waters, Milford, USA) coupled to a photodiode array (PDA) detector (Waters, Milford, USA) was used for the quantification of VL and GUA concentrations. The drawn solutions were first filtered through a 0.2 μm Chromafil ®Xtra PTFE filter (Macherey-Nagel GmbH & Co. KG, Germany). Briefly, the separation of products was performed using an Acquity HSS T3 column

(1.8 μm, 2.1 mm × 100 mm; Waters Corp.). The column oven was held at 30 ºC, and the autosampler was cooled at 4 ºC. The injection volume was set to 5 μL. The binary mobile phase consisted of A (water) and B (acetonitrile). The gradient elution was performed at a flow rate of 0.2 mL/min: 0–1 min, 10% eluent B; 1–25 min, linear increase to 90% eluent B; 25–29.9 min, hold 90% eluent B; 29.9–30 min, decrease to 10% eluent B; 30–35 min, re-equilibrate at 10% eluent B for 5 min. Standard solutions of VL and GUA ranging from 10 to 130 μM were analyzed along with samples and blanks using the channels with UV absorption at 300 and 274 nm, respectively. The calibration curves for VL and GUA standard solutions are shown in Figure S2.

Text S4. IC analyses of small organic acids.

The small organic acids were analyzed using an ion chromatography system (IC, Dionex ICS-1100, Sunnyvale, CA) equipped with a Dionex AS-DV autosampler (Sunnyvale, CA). The separation was achieved using an IonPac™ AS15 column (4 × 250 mm) with an AG15 guard column (4 × 50 mm). The isocratic gradient was applied at a flow rate of 1.2 mL/min with 38 mM sodium hydroxide (NaOH) as the eluent. The total run time was set at 20 min. The standard solutions (1–50 μM) of formic, succinic, and oxalic acid were analyzed three times along with the samples and water blank. Formic, succinic, and oxalic acid had retention times of 3.6 min, 8.3 min, and 11.9 min, respectively.

Text S5. UHPLC-qToF-MS analyses.

The characterization of reaction products was performed using a UHPLC system (ExionLC™ AD, ABSciex, Concord, Canada) coupled to a quadrupole time-of-flight mass spectrometer (qToF-MS) (TripleTOF 6600+, ABSciex). The settings (e.g., column, mobile phase, gradient, oven temperature) in the UHPLC system were the same as those used in UHPLC-PDA

(Text S3). The mass spectrometer was equipped with an electrospray ionization (ESI) source and operated in the positive ion mode (the negative ion mode signals were too low for our analyses) at a resolving power (full width at half-maximum (fwhm) at m/z 300) of 30000 in MS and 30000 in MS/MS (high-resolution mode). Information-dependent acquisition (IDA) scanning was adapted for product identification. The acquisition using IDA consisted of a ToF-MS scan and information-dependent trigger events. The ToF-MS scan had an accumulation time of 250 ms and covered a mass range of m/z 30–700 with a declustering potential (DP) of 40 and collision energy (CE) of 10 eV. The accumulation time for the IDA experiment was 100 ms, and the MS/MS scan range was set from m/z 30–700 in high-resolution mode. The IDA criteria were as follows: 5 most intense ions (number of IDA experiments) with an intensity threshold above 50 cps, isotope exclusion was switched off, and dynamic background subtraction was switched on. The automated calibration device system (CDS) was set to perform an external calibration every four samples. The ESI source conditions were as follows: temperature, 500 °C; curtain gas (CUR), 25 psi; ion source gas 1 at 50 psi; ion source gas 2 at 50 psi; and ion-spray voltage floating (ISVF) at 4.5 kV.

All parameters in the liquid chromatography system and mass spectrometer were controlled using Analyst TF Software 1.8 (ABSciex). The high-resolution LC-MS data were processed with PeakView and Analyst in the SCIEX OS software 1.5 (ABSciex). Peaks from the blank sample were subtracted from the sample signals. In addition to a minimum signal-to-noise ratio of 30, a peak was determined as a product if the difference in peak area between the samples before and after irradiation is ≥ 10 times. The formula assignments were carried out using the MIDAS molecular formula calculator (http://magnet.fsu.edu/~midas/) with the following constraints: C ≤ 35, H ≤ 70, N ≤ 5, O ≤ 20, Na ≤ 1, and the mass error was initially set as 10 ppm. The nitrogen atom was removed in the constraints for the experiments without AN or SN. The detected adducts

in ESI positive ion mode have several types (e.g., $[M+H]^+$, $[M+Na]^+$), and their formation can be influenced by the sample matrix (Erngren et al., 2019). For simplification purposes, we mainly considered $[M+H]^+$ adducts for formula assignments, except for specific experiments with AN or SN in which $[M+NH_4]^+$ adducts and $[M+Na]^+$ adducts were observed. The final assigned formulas were constrained by a mass error mostly < 5 ppm, which is a requirement for product identification using positive ion mode (Roemmelt et al., 2015). The double bond equivalent (DBE) values and carbon oxidation state (OS_c) of the neutral formulas were calculated using the following equations (Koch and Dittmar, 2006):

$$DBE = C - H/2 + N/2 + 1 \quad \text{(Eq. S1)}$$

$$OS_c = 2 \times O/C - H/C \quad \text{(Eq. S2)}$$

where C, H, O, and N correspond to the number of carbon, hydrogen, oxygen, and nitrogen atoms in the neutral formula, respectively. Based on the identified products, the average oxygen to carbon (O:C) ratios, <O:C>: ($< \mathrm{O{:}C} > = \sum_i(\mathrm{abundance_i})O_i / \sum_i(\mathrm{abundance}_i)C_i$) and average hydrogen to carbon (H:C) ratios, <H:C>: ($< \mathrm{H{:}C} > = \sum_i(\mathrm{abundance}_i)H_i / \sum_i(\mathrm{abundance_i})C_i$) after the reactions were further estimated using the signal-weighted method (Bateman et al., 2012). The average OS_c ($<OS_c>$) was calculated as follows:

$$<OS_c> = 2 \times <O:C> - <H:C> \quad \text{(Eq. S3)}$$

Based on the typical MS/MS fragmentation behavior for individual functional groups (Table S1) and DBE values, examples of structures for products identified from VL (and GUA) photo-oxidation experiments were proposed (Table S2~~3~~).

Text S6. Photon flux measurements.

In this work, 2-nitrobenzaldehyde (2NB), a chemical actinometer, was used to determine the photon flux in the aqueous photoreactor. ~~We first measured the relative intensity of light passing through the empty reactor, then the reactor containing 50 μM 2NB using a high-sensitivity spectrophotometer (Brolight Technology Co. Ltd, Hangzhou, China) equipped with an optical fiber (Brolight). Then, the average relative intensity absorbed by 2NB solution as a function of wavelength was calculated.~~ Briefly, the photolysis of 50 μM 2NB in the reactor was monitored by determining its concentration every 5 min for a total of 35 min, during which 2NB was almost completely decayed. The concentration of 2NB was measured using UHPLC-PDA, and the settings (e.g., column, mobile phase, gradient, oven temperature) were the same as those for VL decay analysis (Text S3). The channel with UV absorption at 254 nm was used for the quantification of 2NB. The concentration of 2NB in the reactor followed exponential decay, and its decay rate constant was determined using the following equation:

$$ln\left(\frac{[2\mathrm{NB}]_t}{[2\mathrm{NB}]_0}\right) = -j(2\mathrm{NB}) \times t \qquad \text{(Eq. S4)}$$

where $[2\mathrm{NB}]_t$ and $[2\mathrm{NB}]_0$ are the 2NB concentrations at time t and 0, respectively. The calculated 2NB decay rate constant, j(2NB), was 0.0026 s^{-1}. The following equation can also be used to calculate j(2NB):

$$j(2\mathrm{NB}) = 2.303 \times (10^3\ \mathrm{cm}^3\ \mathrm{L}^{-1} \times\ 1\ \mathrm{mol/N_A\ mlc}) \times \sum\left(I_\lambda^{'} \times \Delta\lambda \times \varepsilon_{2\mathrm{NB},\lambda} \times \Phi_{2\mathrm{NB}}\right) \qquad \text{(Eq. S5)}$$

where N_A is Avogadro's number, $I_\lambda^{'}$ is the actinic flux (photons cm^{-2} s^{-1} nm^{-1}), $\Delta\lambda$ is the wavelength interval between actinic flux data points (nm), and $\varepsilon_{2NB,\lambda}$ and $\Phi_{2NB,\lambda}$ are the base-10 molar absorptivity (M^{-1} cm^{-1}) and quantum yield (molecule $photon^{-1}$) for 2NB, respectively. Values of $\varepsilon_{2NB,\lambda}$ (in water) at each wavelength under 298 K and a wavelength-independent Φ_{2NB} value of 0.41 were adapted from Galbavy et al. (2010). Similar to Smith et al. (2014, 2016), we measured the spectral shape of the photon output of our illumination system (i.e., the relative flux at each wavelength) using a high-sensitivity spectrophotometer (Brolight Technology Co. Ltd, Hangzhou, China). Using a scaling factor (SF), this measured relative photon output, $I_\lambda^{relative}$, is related to I_λ' as follows:

$$I_\lambda' = I_\lambda^{relative} \times SF \qquad \text{(Eq. S6)}$$

Substitution of Eq. S6 into Eq. S5 and rearrangement yields:

$$SF = \frac{j(2NB)}{2.303 \times (10^3\ cm^3\ L^{-1} \times 1\ mol/N_A\ mlc) \times \sum(I_\lambda^{relative} \times \Delta\lambda \times \varepsilon_{2NB,\lambda} \times \Phi_{2NB})} \qquad \text{(Eq. S7)}$$

and substitution of Eq. S6 into Eq. S7 yields:

$$I_\lambda' = I_\lambda^{relative} \frac{j(2NB)}{2.303 \times (10^3\ cm^3\ L^{-1} \times 1\ mol/N_A\ mlc) \times \sum(I_\lambda^{relative} \times \Delta\lambda \times \varepsilon_{2NB,\lambda} \times \Phi_{2NB})} \qquad \text{(Eq. S8)}$$

Finally, $I_\lambda^{'}$ was estimated through Eq. S8~~5~~. The estimated photon flux in the aqueous reactor is shown in Figure S1~~2~~.

The actinic flux during a haze event over Beijing (40º N, 116º E) on January 12, 2013, at 12:00 pm (GMT+8) (Che et al., 2014) estimated using the National Center for Atmospheric Research Tropospheric Ultraviolet-Visible (TUV) Radiation Model

(http://cprm.acom.ucar.edu/Models/TUV/Interactive_TUV/) is also shown in Figure S2. The parameters used for the Quick TUV calculator were: Overhead Ozone Column: 300 du; Surface Albedo: 0.1; Ground Elevation: 0 km asl; Measured Altitude: 0 km asl; Clouds optical depth: 0, base: 4, top: 5; Aerosols optical depth: 2.5, single scattering albedo: 0.9, Angstrom exponent: 1; Sunlight direct beam, diffuse down, diffuse up: 1; 4 streams transfer model. For clear days, the actinic flux was estimated over Beijing (at the same date and time) using the default parameters.

Text S7. Estimation of the apparent quantum efficiency of guaiacol photodegradation.

The apparent quantum efficiency of GUA photodegradation (φ_{GUA}) in the presence of either VL or nitrate during simulated sunlight illumination can be defined as (Anastasio et al., 1996; Smith et al., 2014, 2016):

$$\Phi_{GUA} = \frac{\text{mol GUA destroyed}}{\text{mol photons absorbed}} \qquad \text{(Eq. S9)}$$

Φ_{GUA} was calculated using the measured rate of GUA decay and rate of light absorption by either VL or nitrate through the following equation:

$$\Phi_{GUA} = \frac{\text{rate of GUA decay}}{\text{rate of light absorption by VL or nitrate}} = \frac{k'_{GUA} \times [GUA]}{\sum[\left(1-10^{-\varepsilon_\lambda [C] l}\right) \times I'_\lambda]} \qquad \text{(Eq. S10)}$$

where k'_{GUA} is the pseudo-first-order rate constant for GUA decay, [GUA] is the concentration of GUA (M), ε_λ is the base-10 molar absorptivity (M^{-1} cm^{-1}) of VL or nitrate at wavelength λ, [C] is the concentration of VL or nitrate (M), l is the pathlength of the illumination cell (cm), and I'_λ is the volume-averaged photon flux (mol-photons L^{-1} s^{-1} nm^{-1}) determined from 2NB actinometry:

$$j(2NB) = 2.303 \times \Phi_{2NB} \times l \times \Sigma_{300\,nm}^{350\,nm} (\varepsilon_{2NB,\lambda} \times I'_\lambda \times \Delta\lambda) \qquad \text{(Eq. S11)}$$

Table S1. Typical fragmentation behavior observed in MS/MS spectra for individual functional groups from Holčapek et al. (2010).

Functional group	**Fragment ions**	**MS/MS loss**
Nitro (RNO_2)	$[M+H-OH]^{+\bullet}$	-OH
	$[M+H-H_2O]^{+}$	$-H_2O$
	$[M+H-NO]^{+\bullet}$	-NO
	$[M+H-NO_2]^{+\bullet}$	$-NO_2$
Nitroso (RNO)	$[M+H-NO]^{+}$	-NO
Carboxylic acid (ROOH)	$[M+H-H_2O]^{+}$	$-H_2O$
	$[M+H-CO_2]^{+}$	$-CO_2$
	$[M+H-H_2O-CO]^{+}$	$-H_2O-CO$
Phenol (ROH)	$[M+H-H_2O]^{+}$	$-H_2O$
	$[M+H-CO]^{+}$	-CO
Methoxy ($ROCH_3$)	$[M+H-CH_3]^{+\bullet}$	$-CH_3$
	$[M+H-CH_3O]^{+\bullet}$	$-CH_3O$
	$[M+H-CH_3OH]^{+}$	$-CH_3OH$
	$[M+H-HCOH]^{+}$	-HCOH
Ester (R^1COOR^2)	$[M+H-R^2OH]^{+}$	$-R^2OH$
	$[M+H-R^2OH-CO]^{+}$	$-R^2OH-CO$
Amine	$[M+H-NH_3]^{+}$	$-NH_3$
Aldehyde (RCHO)	$[M+H-CO]^{+}$	-CO

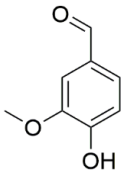
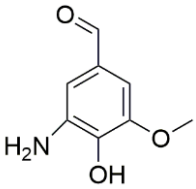
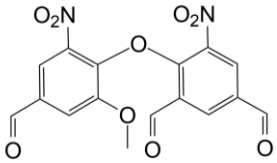
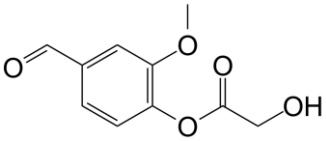
Table S2. Reaction conditions, initial VL (and GUA) decay rates, normalized abundance of products, and average carbon oxidation state (<OS_c>) in each experiment. Except where noted, the reaction systems consisted of VL (0.1 mM); GUA (0.1 mM), AN (1 mM); sodium nitrate (SN) (1 mM); VOC (IPA) (1 mM) or inorganic anions (NaBC) (1 mM) under air-saturated conditions after 6 h of simulated sunlight irradiation. Analyses were performed using UHPLC-qToF-MS equipped with an ESI source and operated in the positive ion mode.

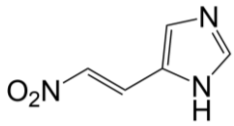
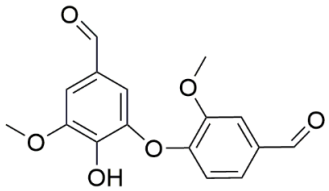
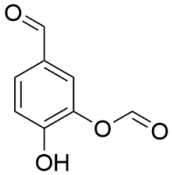
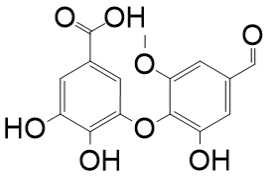
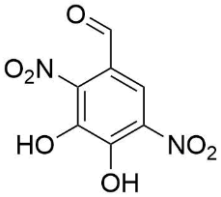
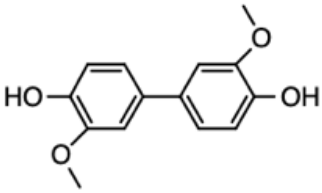
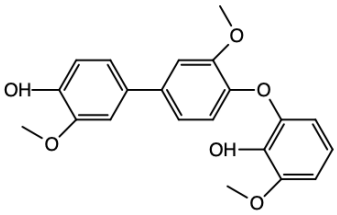
Exp no.	pH	Reaction conditions	Initial VL (and GUA) decay rates (min^{-1})[b]	Ratio of 50 most abundant products to total products[c]	Normalized abundance of products[d]	Normalized abundance of N-containing compounds[d]	<OS_c>[e]
A1	2.5	VL*	2.8×10^{-3}	0.59	1.7	-	-0.05
A2		VL+AN	2.5×10^{-3}	0.63	1.4	5.3×10^{-2}	-0.04
A3	3	VL*	2.1×10^{-3}	0.53	1.9	--	-0.04
A4		VL+AN	2.1×10^{-3}	0.56	1.9	3.6×10^{-2}	-0.05
A5	4	VL*	1.9×10^{-3}	0.58	0.26	--	-0.16
A6		VL* (N_2-saturated)	4.6×10^{-4}	0.96	4.7×10^{-2}	--	-0.24
A7		VL+AN	1.9×10^{-3}	0.53	0.37	1.7×10^{-2}	-0.13
A8		VL+AN (N_2-saturated)	2.9×10^{-4}	0.89	0.12	6.3×10^{-3}	-0.21
A13		VL+SN	--	0.51	0.42	1.7×10^{-2}	-0.07
A14		VL* (0.01 mM)[a]	--	0.90	0.37	--	-0.07

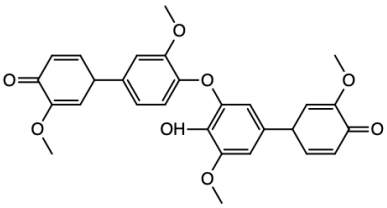
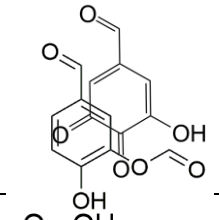
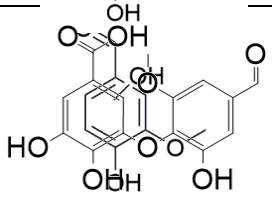
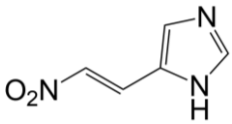
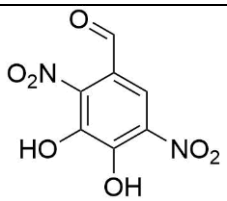
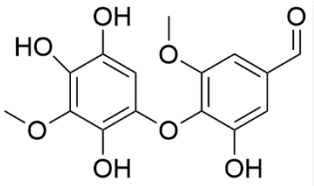
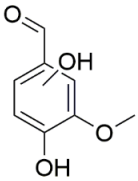
~~A15~~		~~VL (0.01 mM) + AN (0.01 mM)~~	~~–~~	~~0.77~~	~~0.40~~	~~8.6×10^{-3}~~	~~0.12~~
~~A16~~		~~VL (0.01 mM) + AN~~	~~–~~	~~0.42~~	~~0.45~~	~~1.2×10^{-2}~~	~~0.06~~
~~A17~~		~~GUA only~~	~~9.0×10^{-4}~~	~~0.77~~	~~–~~	~~–~~	~~0.28~~
~~A18~~		~~GUA+VL~~	~~GUA: 2.0×10^{-3} VL: 6.2×10^{-4}~~	~~0.60~~	~~2.2~~	~~–~~	~~0.27~~
~~A19~~		~~GUA+AN~~	~~1.1×10^{-3}~~	~~0.77~~	~~–~~	~~–~~	~~0.26~~

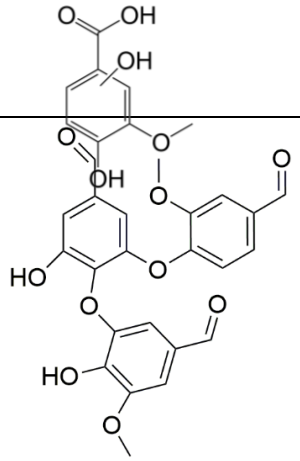
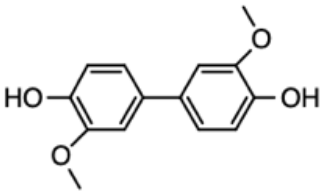
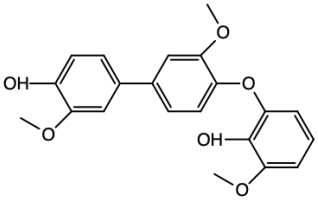
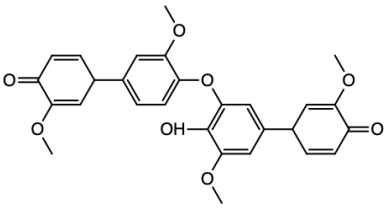
~~[a]Irradiation time for VL* (0.01 mM, A14) was 3 h. [b]The data fitting was performed in the initial linear region. [c]Ratio of the normalized abundance of the 50 most abundant products to that of total products, except for direct GUA photodegradation (A17), GUA+VL (A18), and GUA+AN (A19) whose ratios are based on the absolute signal area of products. [d]The normalized abundance of products was calculated using Eq. 2. [e]<OS_c> of the 50 most abundant products.~~

Table S2~~3~~. Examples of proposed molecular structures for products identified from vanillin (and guaiacol) photo-oxidation experiments in this study.

No.	Formula	DBE	Proposed structure	MS/MS fragment ions		
1	$C_8H_8O_3$ (VL; the aqSOA precursor)	5		-CO-CH_3OH	-CO	-CO-CH_3OH-CO
2	$C_8H_9NO_3$	5		-CO-CH_3	-NH_3	
3~~3~~	$C_{16}H_{10}N_2O_9$	13		-NO_2		
4	$C_{10}H_{10}O_5$	6		-CH_3OH	-CH_3OH-CO	
~~4~~	~~$C_{10}H_{10}O_5$~~	~~6~~		~~-CH_3OH~~	~~-CH_3OH-CO~~	

5	$C_5H_5N_3O_2$	5		-NH		
6~~5~~	$C_{16}H_{14}O_6$	10		-CO-CH_3OH-CO	-CO-CH_3OH-CO-CH_3OH	-CO-CH_3OH-CO-CO
7	$C_8H_6O_4$	6		-CO	-CO-CO	
8	$C_{15}H_{12}O_8$	10		-COOH		
9	$C_7H_4N_2O_7$	7				
10	$C_{14}H_{14}O_4$	8				
11~~6~~	$C_{21}H_{20}O_6$~~$C_7H_4O_4$~~	12 ~~6~~		~~-CO~~	~~-CO-CO~~	

127	$C_{28}H_{24}O_8$~~$C_{15}H_{12}O_8$~~	17~~10~~		~~-CO~~	~~-CH_3~~	
138	$C_7H_4O_4$~~$C_8H_6O_4$~~	6 ~~6~~		-CO~~-CO~~	-CO-CO-~~CO-CO~~	
149	$C_8H_8O_4$~~$C_{15}H_{12}O_8$~~	5~~10~~		-CO-CH_3OH-~~COOH~~	-CO	-H_2O
~~10~~	~~$C_5H_5N_3O_2$~~	~~5~~		~~-NH~~		
~~11~~	~~$C_7H_4N_2O_7$~~	~~7~~				
~~12~~	~~$C_{15}H_{14}O_8$~~	~~9~~		~~-CO-CH_3OH-CO-CO~~	~~-CO-CH_3OH-CO-H_2O~~	~~-CO-CH_3OH-CO~~
~~13~~	~~$C_8H_8O_4$~~	~~5~~		~~-CO-CH_3OH~~	~~-CO~~	~~-H_2O~~
~~14~~	~~$C_8H_8O_5$~~	~~5~~		~~-CO-CH_3OH~~	~~-CO_2~~	~~-CH_3OH~~

~~15~~	~~$C_{23}H_{18}O_9$~~	~~15~~				
~~19~~	~~$C_{14}H_{14}O_4$~~	~~8~~				
~~20~~	~~$C_{21}H_{20}O_6$~~	~~12~~				
~~21~~	~~$C_{28}H_{24}O_8$~~	~~17~~				

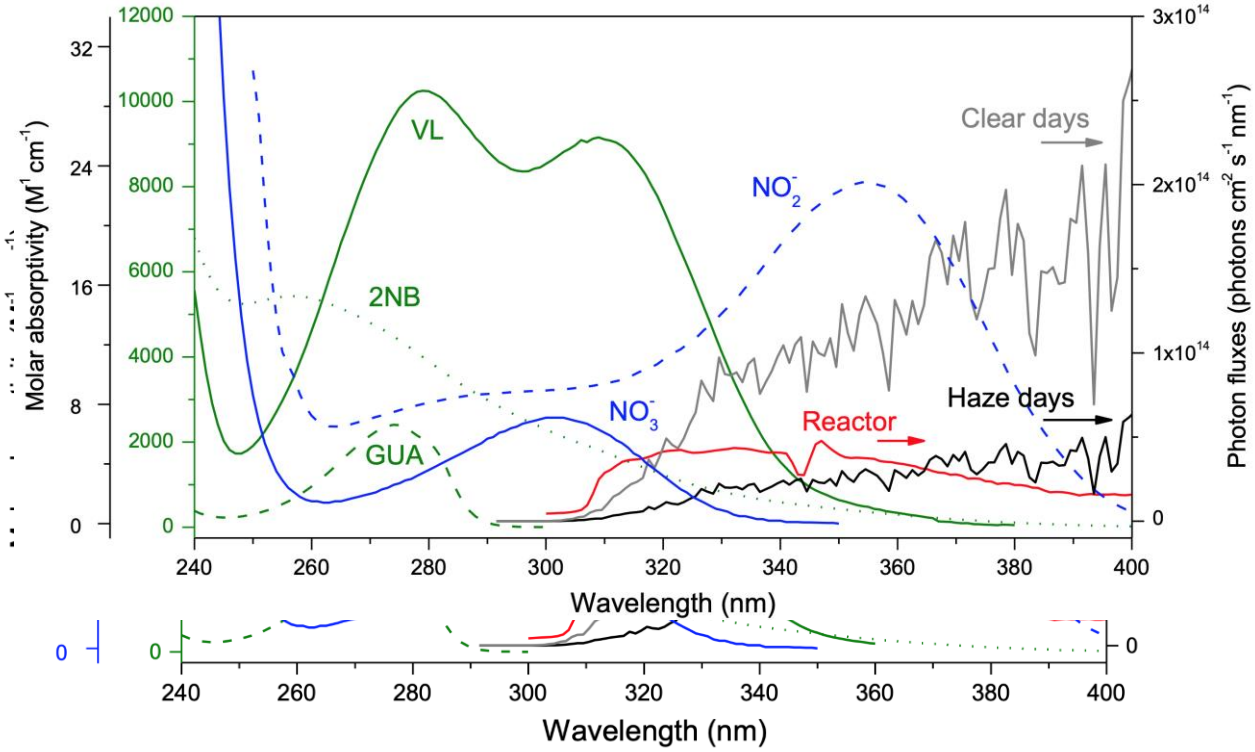


Figure S1. The base-10 molar absorptivities (ε, M^{-1} cm^{-1}) of vanillin (VL, green solid line), 2-nitrobenzaldehyde (2NB, green dotted line), guaiacol (GUA, green dashed line), NO_2^- (blue dashed line), NO_3^- (blue solid line), and photon flux in the aqueous reactor (red line) during typical haze days (black line) or clear days (grey line) in Beijing, China. The ε values for 2NB and NO_2^- were adapted from Galbavy et al. (2010) and Chu and Anastasio (2007), respectively.

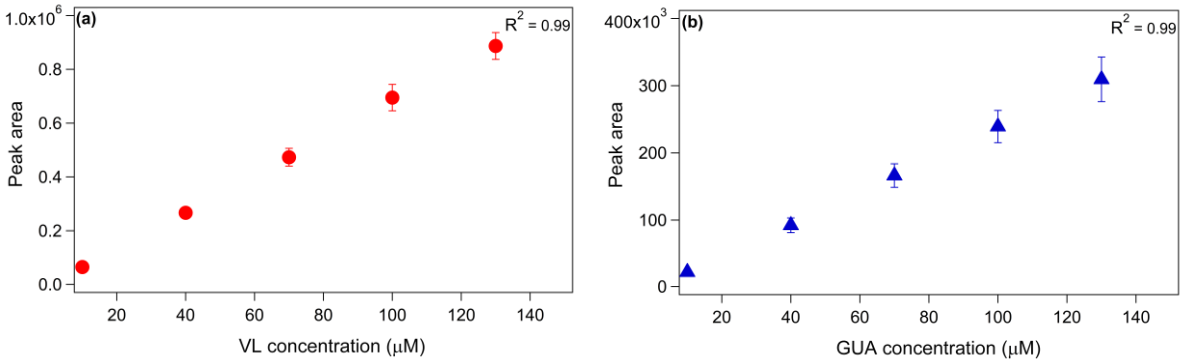


Figure S2. Calibration curves for (a) VL and (b) GUA standard solutions (10–130 μM). Error bars represent one standard deviation.

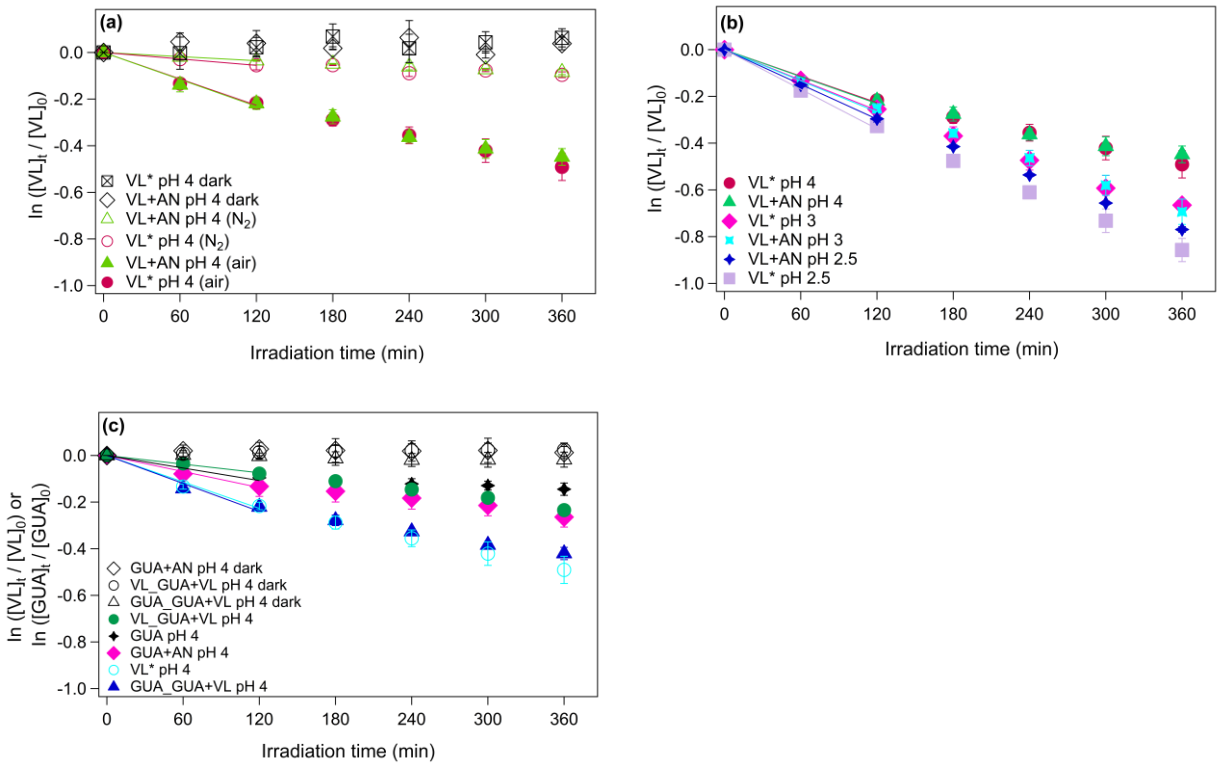


Figure S3. (a–c~~d~~) The decay of VL under different experimental conditions for direct photosensitized oxidation of VL (VL*) and nitrate-mediated VL photo-oxidation (VL+AN): (a) ~~Effect of secondary oxidants from VL triplets on~~ VL* and VL+AN at pH 4 under N_2- (A6, A8) and air-saturated (A5, A7) conditions. (b) Effect of pH on VL* and VL+AN at pH 2.5 (A1, A2), 3 (A3, A4), and 4 (A5, A7) under air-saturated conditions. ~~(c) Effect of the presence of VOCs and inorganic anions: IPA (A9) and NaBC (A10) on VL* at pH 4 under air-saturated conditions. (d) Effect of the presence of VOCs and inorganic anions: IPA (A11) and NaBC (A12) on VL+AN at pH 4 under air-saturated conditions.~~ (c~~e~~) The decay of VL (and GUA) during direct GUA photodegradation (A13~~7~~) and photo-oxidation of GUA in the presence of VL (GUA+VL; A14~~8~~) or nitrate (GUA+AN; A15~~9~~) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. Error bars represent one~~1~~ standard deviation; most error bars are smaller than the markers.

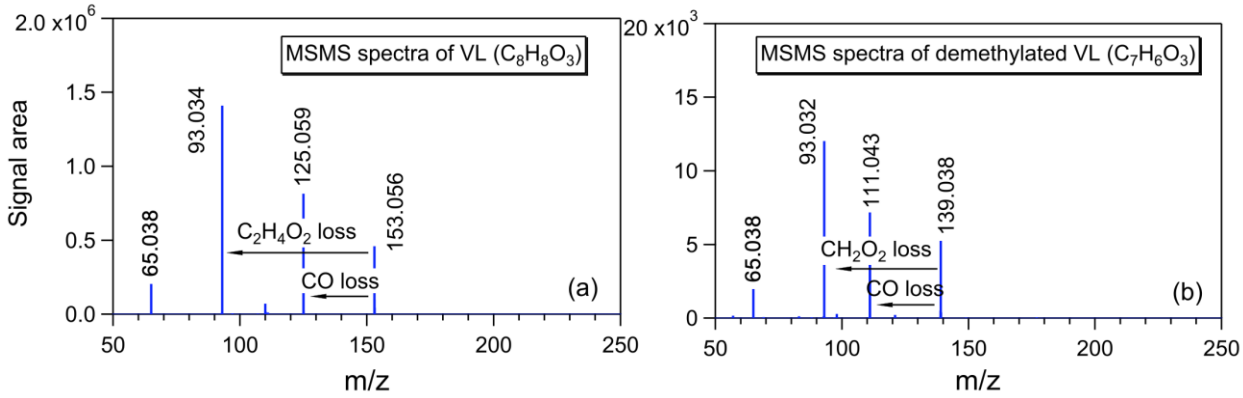


Figure S4. MS/MS spectra of (a) VL and (b) demethylated VL. The arrows indicate possible fragmentation pathways of VL and demethylated VL.

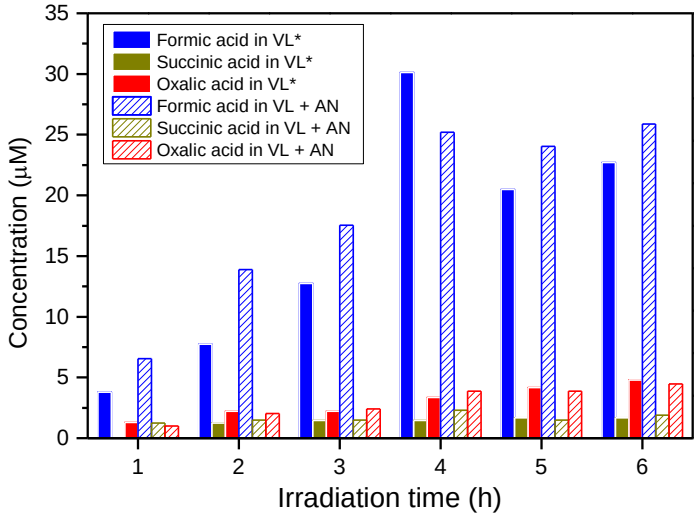


Figure S5. The concentration of formic, oxalic, and succinic acid at different reaction times for VL* (A5) and VL+AN (A7) at pH 4 under air-saturated conditions.

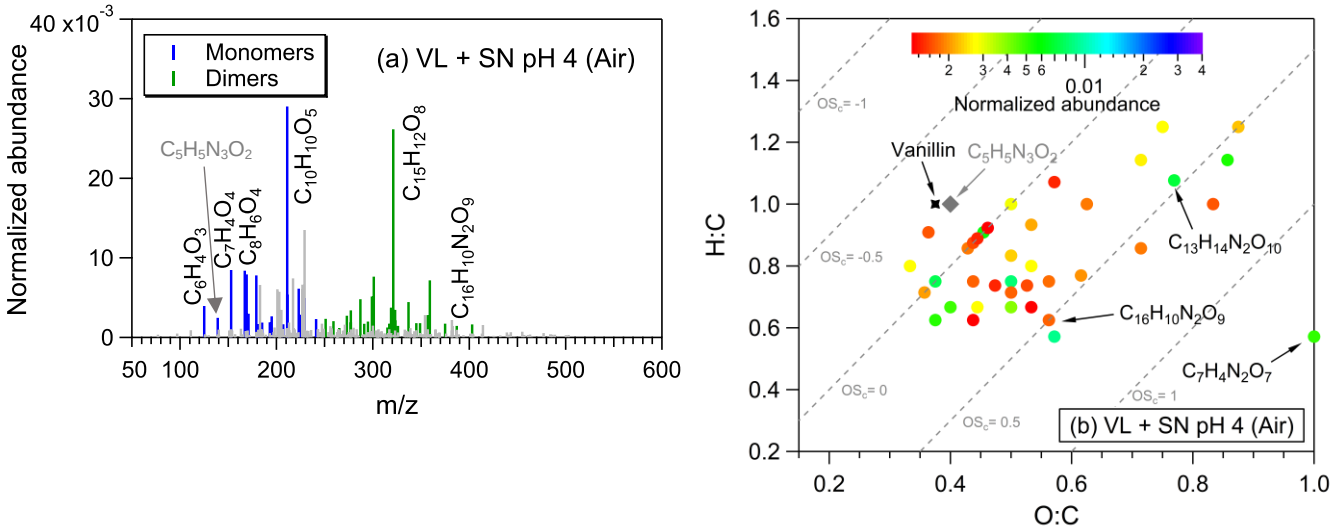


Figure S6. (a) Reconstructed mass spectra of assigned peaks and (b) van Krevelen diagram of the 50 most abundant products from VL+SN (A9~~13~~) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. The normalized abundance of products was calculated from the ratio of the peak area of the product to that of VL (Eq. 2). The 50 most abundant products contributed more than half of the total normalized abundance of products, and they were identified as monomers (blue) and dimers (green). Grey peaks denote peaks with low abundance or unassigned formula. Examples of high-intensity peaks were labeled with the corresponding neutral formulas. The color bar denotes the normalized abundance of products. The grey dashed lines indicate the carbon oxidation state values (e.g., OS_c = -1, 0, and 1). The grey arrows show where the potential imidazole derivative ($C_5H_5N_3O_2$) from VL+AN was observed.

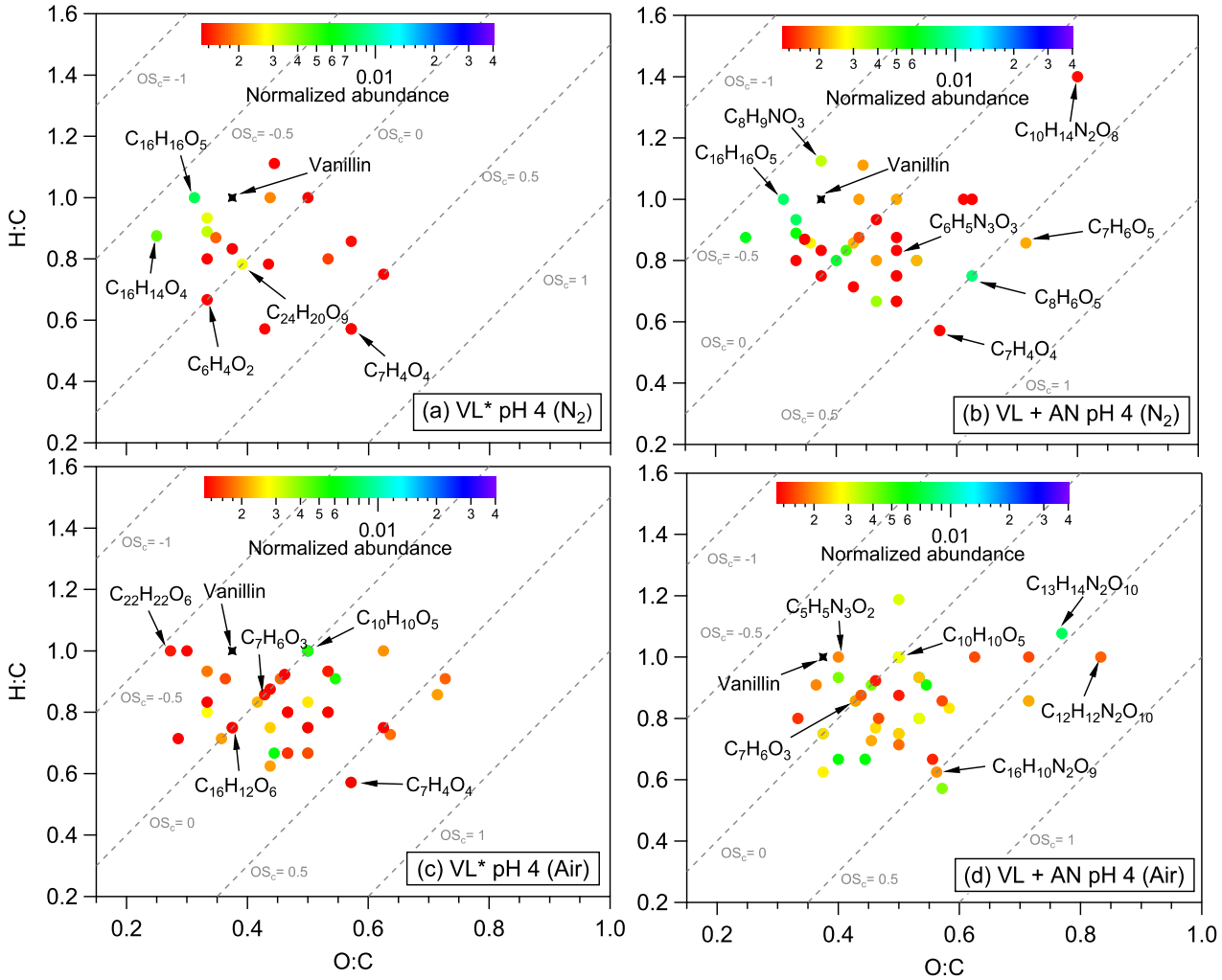


Figure S7. van Krevelen diagrams of the 50 most abundant products from (a) VL* (N_2-saturated; A6), (b) VL+AN (N_2-saturated; A8), (c) VL* (air-saturated; A5), and (d) VL+AN (air-saturated; A7) at pH 4 after 6 h of simulated sunlight irradiation. The color bar denotes the normalized abundance of products. The grey dashed lines indicate the carbon oxidation state values (e.g., OS_c = -1, 0, and 1).

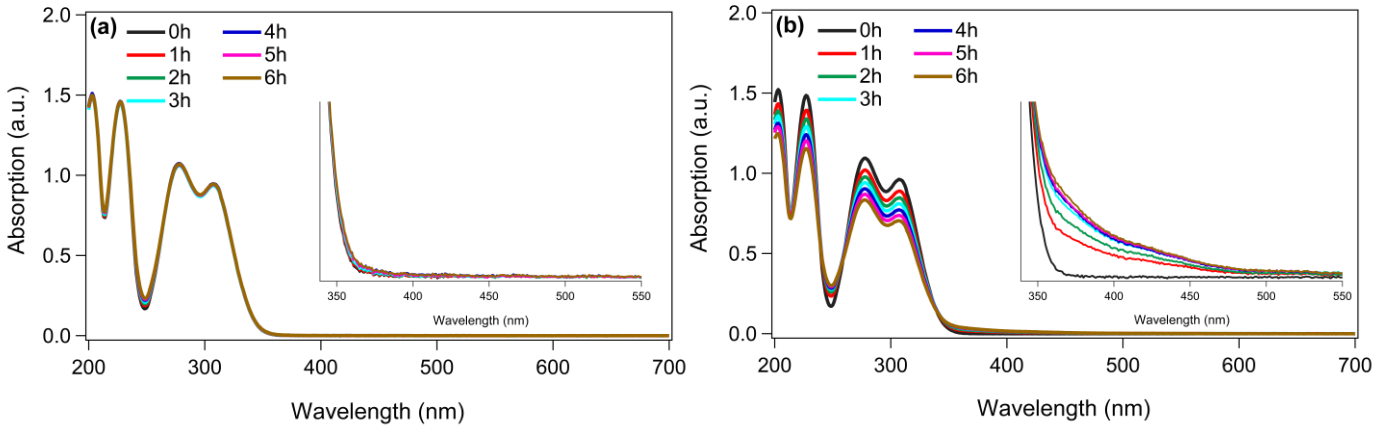


Figure S8. UV–Vis absorption spectra of VL* (A6, A5; pH 4) under (a) N_2- and (b) air-saturated conditions at different time intervals. The insets show the absorbance enhancement from 350 to 550 nm.

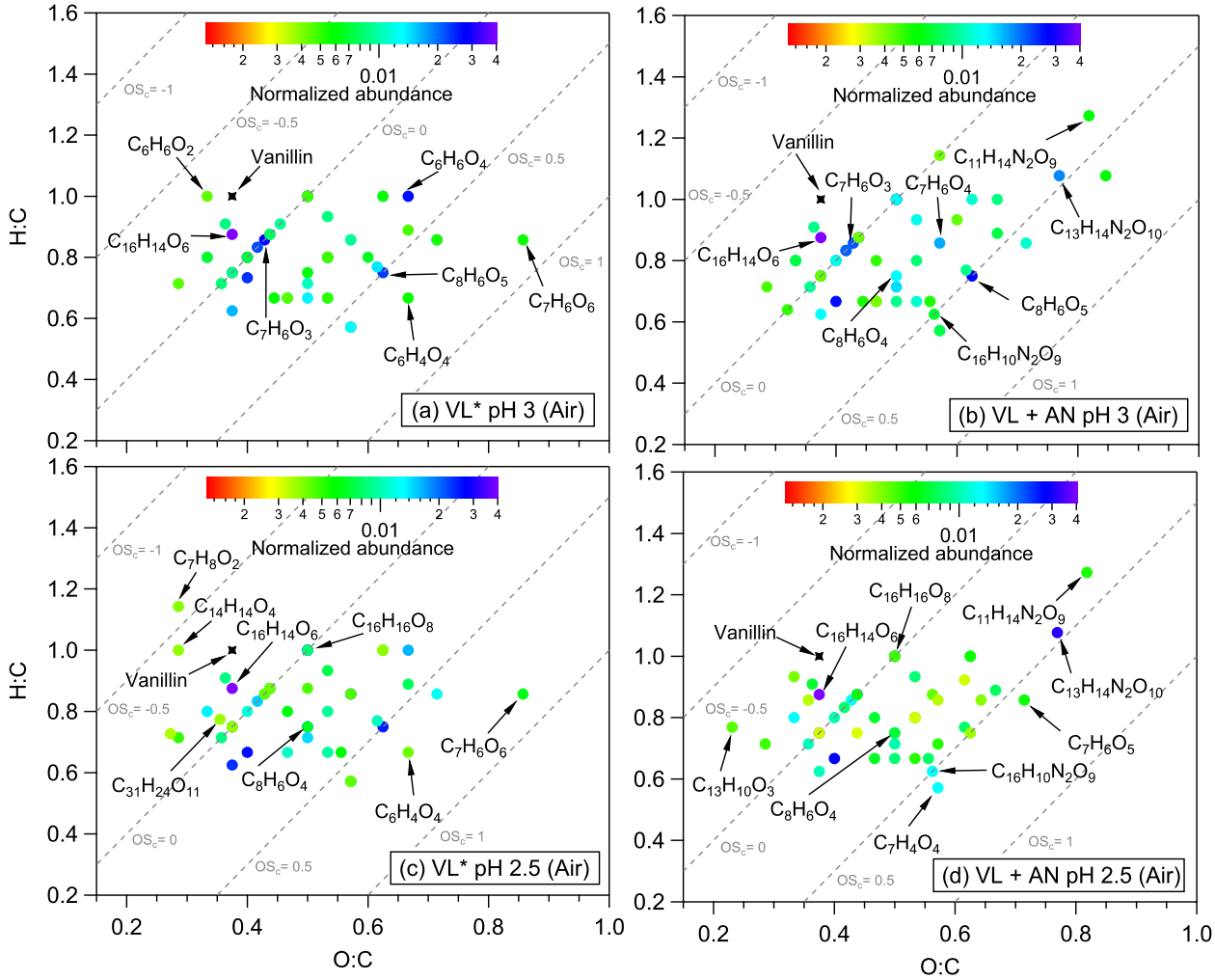


Figure S9. van Krevelen diagrams of the 50 most abundant products from (a) VL* pH 3 (A3), (b) VL+AN pH 3 (A4), (c) VL* pH 2.5 (A1), and (d) VL+AN pH 2.5 (A2) under air-saturated conditions after 6 h of simulated sunlight irradiation. The color bar denotes the normalized abundance of products. The grey dashed lines indicate the carbon oxidation state values (e.g., OS_c =-1, 0, and 1).

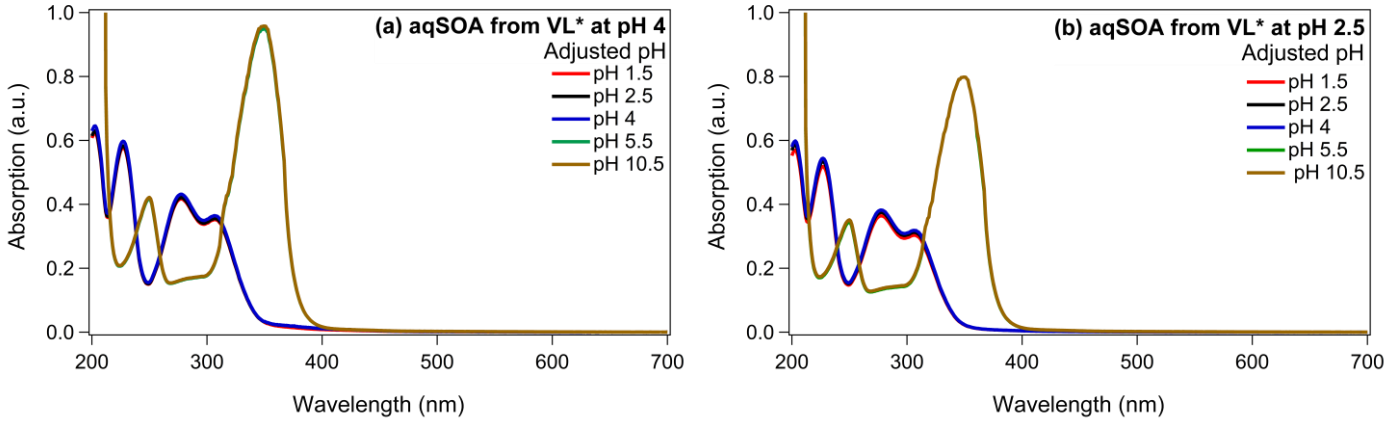


Figure S10. UV–Vis absorption spectra of VL*-derived aqSOA formed at (a) pH 4 and (b) pH 2.5 over a range of pH conditions from 1.5 to 10.5.

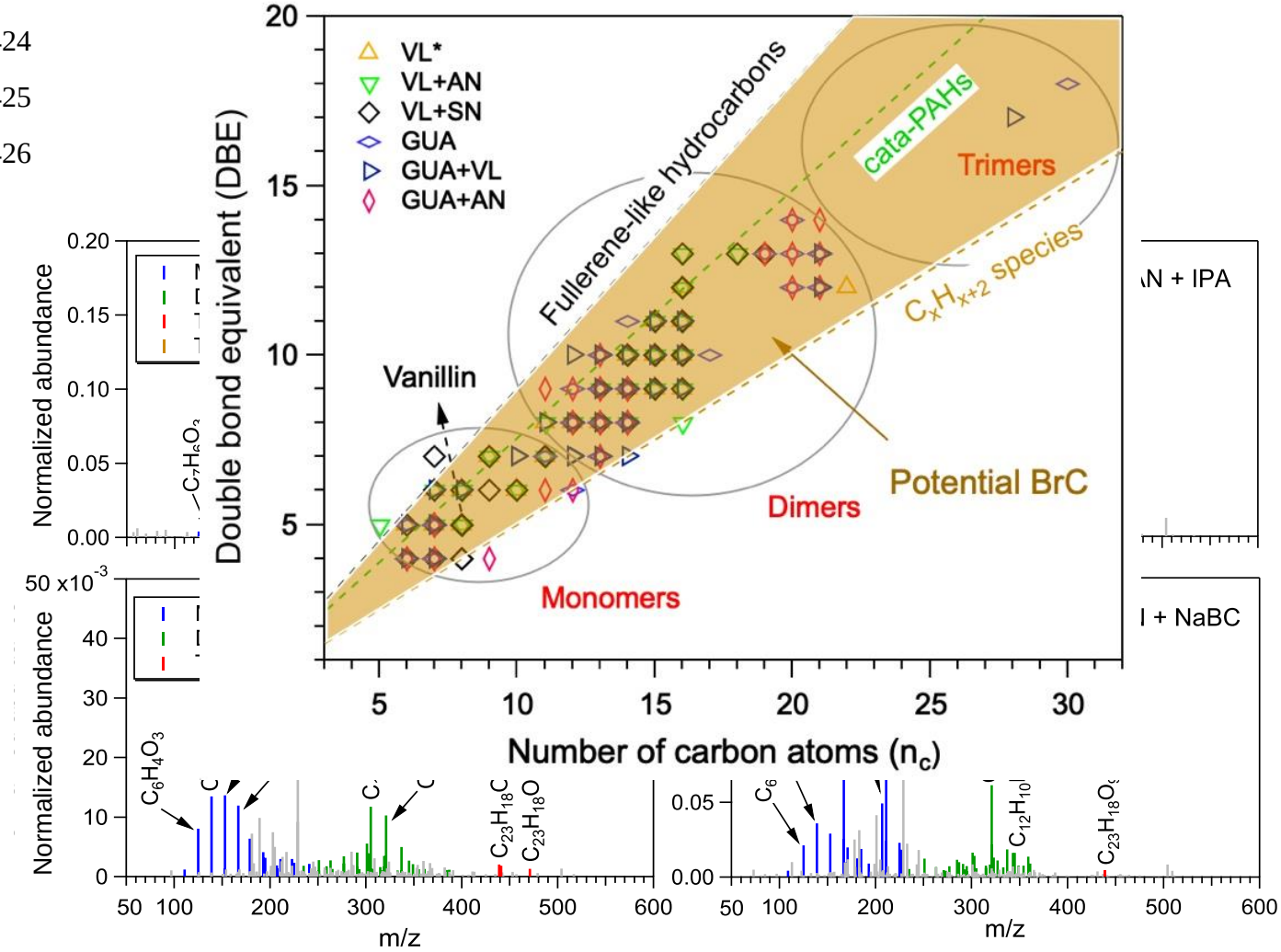
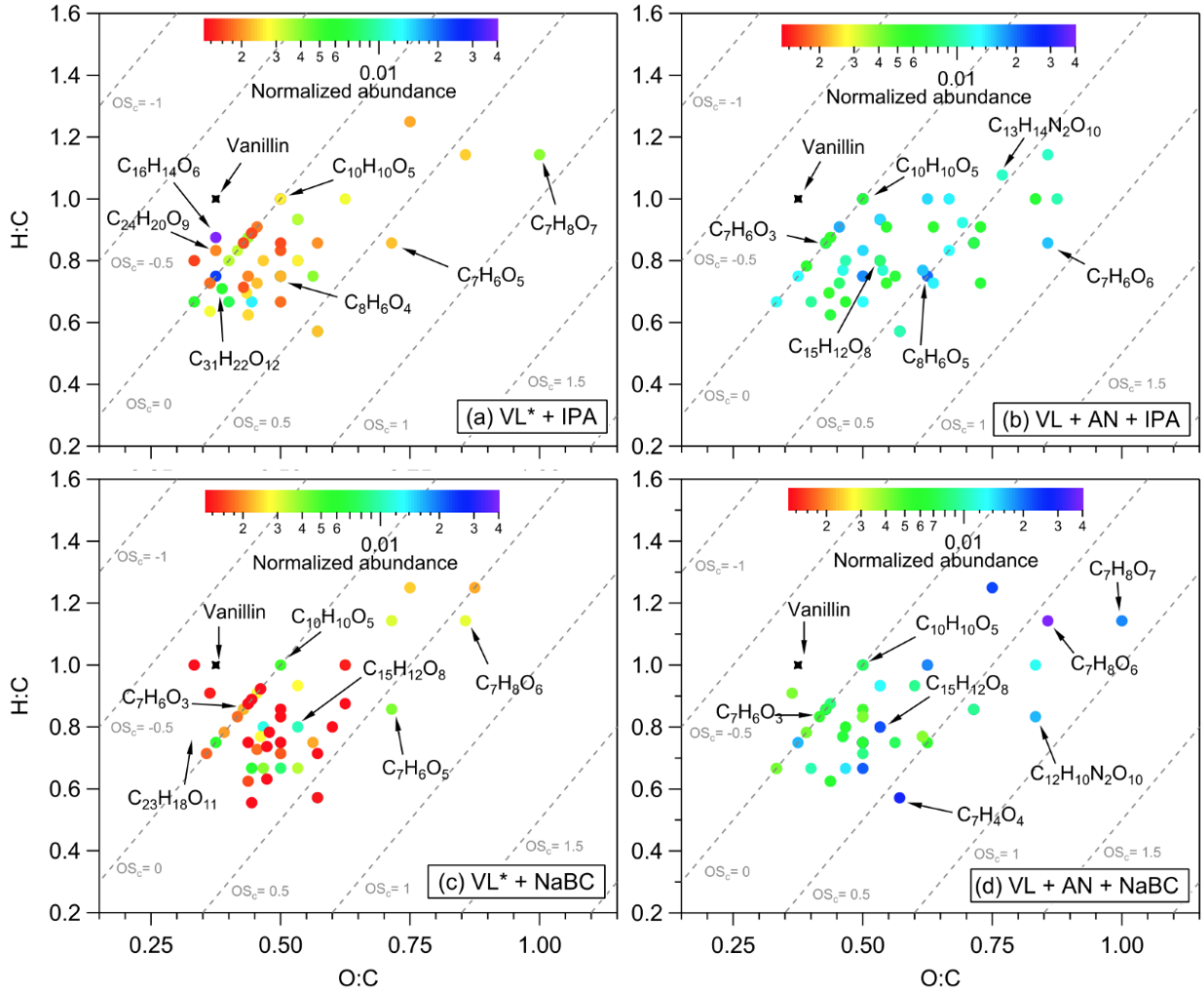


Figure S10. Reconstructed mass spectra of assigned peaks from (a) VL*+IPA (A9), (b) VL+AN+IPA (A11), (c) VL*+NaBC (A10), and (d) VL+AN+NaBC (A12) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. The normalized abundance of products was calculated from the ratio of the peak area of the product to that of VL (Eq. 2). The 50 most abundant products contributed more than half of the total normalized abundance of products, and they were identified as monomers (blue), dimers (green), trimers (red), and tetramers (orange). Grey peaks denote peaks with low abundance or unassigned formula. Examples of high-intensity peaks were labeled with the corresponding neutral formulas. Note the different scales on the y-axes.



~~**Figure S11.** van Krevelen diagrams of the 50 most abundant products from (a) VL*+IPA (A9), (b) VL+AN+IPA (A11), (c) VL*+NaBC (A10), and (d) VL+AN+NaBC (A12) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. The color bar denotes the normalized abundance of products. The grey dashed lines indicate the carbon oxidation state values (e.g., OS_c = -1, 0, and 1).~~

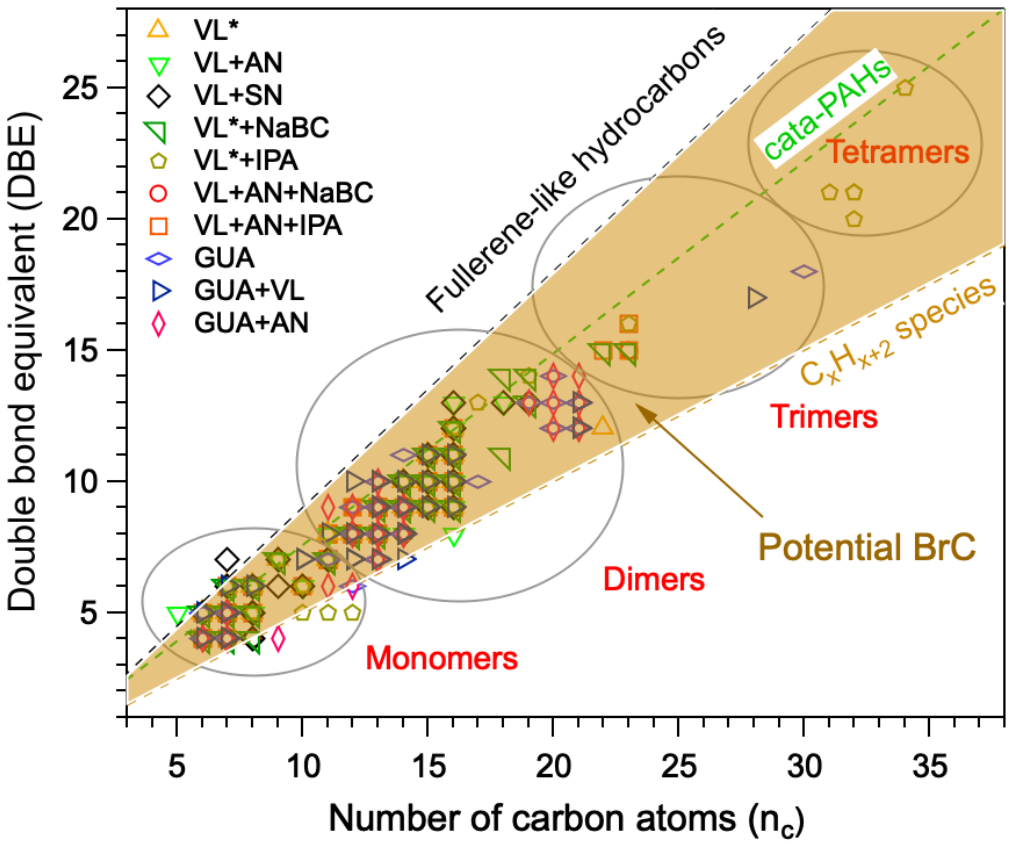


Figure S11~~2~~. The plot of the double bond equivalent (DBE) values vs. number of carbon atoms (n_c) (Lin et al., 2018) for the 50 most abundant products from pH 4 experiments under air-saturated conditions. Dashed lines indicate DBE reference values of fullerene-like hydrocarbons (Lobodin et al, 2012) (black dashed line), cata-cond~~s~~ensed polycyclic aromatic hydrocarbons (PAHs) (Siegmann and Sattler, 2000) (green dashed line), and linear conjugated polyenes (general formula C_xH_{x+2}) (brown dashed line). Data points within the shaded area are potential BrC chromophores.

Light grey circles show the classification of the data points as monomers, dimers, trimers, or tetramers.

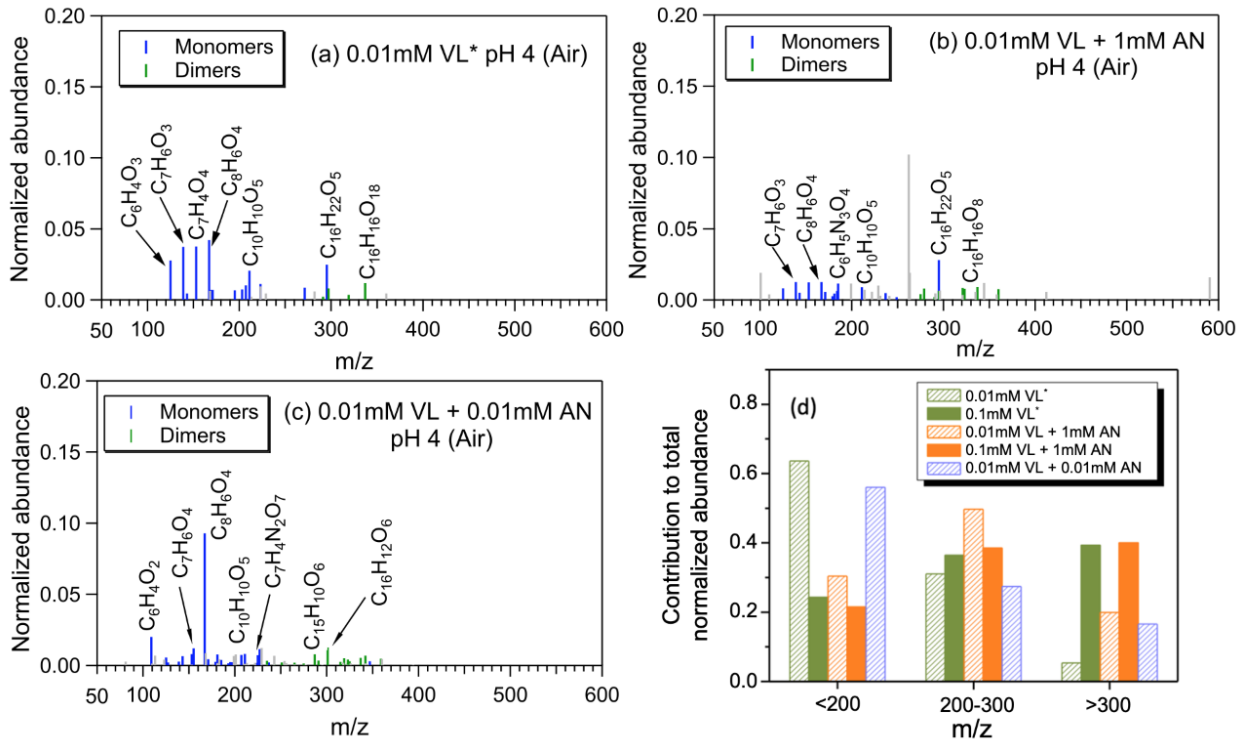


Figure S12~~3~~. Reconstructed mass spectra of assigned peaks from (a) 0.01 mM VL* (A10~~4~~), (b) 0.01 mM VL + 1 mM AN (A11~~6~~), and (c) 0.01 mM VL + 0.01 mM AN (A12~~5~~) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation. The normalized abundance of products was calculated from the ratio of the peak area of the product to that of VL (Eq. 2). The 50 most abundant products contributed more than half of the total normalized abundance of products, and they were identified as monomers (blue) and dimers (green). Grey peaks denote peaks with low abundance or unassigned formula. Examples of high-intensity peaks were labeled with the corresponding neutral formulas. (d) Contributions of different m/z ranges to the normalized abundance of products from experiments with low [VL] = 0.01 mM (A10~~4~~–A12~~6~~) and high [VL] = 0.1 mM (A5 and A7) at pH 4 under air-saturated conditions after 6 h of simulated sunlight irradiation.

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