https://doi.org/10.5194/acp-2020-536
Preprint. Discussion started: 7 July 2020

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Mass Accommodation and Gas-Particle Partitioning in Secondary Organic Aerosols: Dependence on Diffusivity, Volatility, Particle-phase Reactions, and Penetration Depth

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https://doi.org/10.5194/acp-2020-536
Preprint. Discussion started: 7 July 2020

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Abstract.

Mass accommodation is an essential process for gas-particle partitioning of organic compounds in secondary organic aerosols (SOA). The mass accommodation coefficient is commonly described as the probability of a gas molecule colliding with the surface to enter the particle phase. It is often applied, however, without specifying if and how deep a molecule has to penetrate beneath the surface to be regarded as incorporated into the condensed phase (adsorption vs. absorption). While this aspect is usually not critical for liquid particles with rapid surface-bulk exchange, it can be important for viscous semisolid or glassy solid particles to distinguish and resolve the kinetics of accommodation at the surface, transfer across the gas-particle interface, and further transport into the particle bulk.

For this purpose, we introduce a novel parameter: an effective mass accommodation coefficient α_{eff} that depends on penetration depth and is a function of surface accommodation coefficient, volatility, bulk diffusivity, and particle-phase reaction rate coefficient. Application of α_{eff} in the traditional Fuchs-Sutugin approximation of mass-transport kinetics at the gas-particle interface yields SOA partitioning results that are consistent with a detailed kinetic multilayer model (KM-GAP, Shiraiwa et al. 2012) and two-film model solutions (MOSAIC, Zaveri et al., 2014) but deviate substantially from earlier modeling approaches not considering the influence of penetration depth and related parameters.

For highly viscous or semisolid particles, we show that the effective mass accommodation coefficient remains similar to the surface accommodation coefficient in case of low-volatile compounds, whereas it can decrease by several orders of magnitude in case of semi-volatile compounds. Such effects can explain apparent inconsistencies between earlier studies deriving mass accommodation coefficients from experimental data or from molecular dynamics simulations.

Our findings challenge the approach of traditional SOA models using the Fuchs-Sutugin approximation of mass transfer kinetics with a fixed mass accommodation coefficient regardless of particle phase state and penetration depth. The effective mass accommodation coefficient introduced in this study provides an efficient new way of accounting for the influence of volatility, diffusivity, and particle-phase reactions on SOA partitioning in process models as well as in regional and global air quality models.

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Introduction.

Secondary organic aerosols (SOA) are major constituents of atmospheric particulate matter, affecting air quality, climate, and public health (Jimenez et al., 2009; Kanakidou et al., 2005; Pöschl and Shiraiwa, 2015; Shrivastava et al., 2017a). Gas-phase reactions of volatile organic compounds (VOC) emitted from various anthropogenic and biogenic sources with oxidants such as ozone and OH radicals lead to the formation and growth of SOA (Kroll and Seinfeld, 2008). The oxidation of VOC forms a myriad of semi-volatile (SVOC) and low volatility organic compounds (LVOC) that can condense on pre-existing particles (Ziemann and Atkinson, 2012) or contribute to nucleation and new particle formation (Tröstl et al., 2016). The evolution of SOA is a complex multi-step process that involves chemical reactions and mass trnsport in the gas phase, at the particle surface and in the particle bulk, but the interplay of these processes and the rate-limiting steps in SOA formation have not yet been fully resolved/elucidated (Shiraiwa et al., 2014).

Traditionally, SOA particles were assumed to be homogeneous and well-mixed quasi-liquid droplets (Pankow, 1994). As demonstrated by recent atmospheric measurements and laboratory experiments, they can adopt glassy solid or amorphous semi-solid phase states, challenging the traditional views of SOA properties, interactions and effects (Koop et al., 2011; Reid et al., 2018; Virtanen et al., 2010). Slow diffusion of water, oxidants and organic molecules in viscous, semi-solid, or glassy particles may lead to kinetic limitations in heterogeneous and multiphase reactions (Alpert et al., 2019; Davies and Wilson, 2015; Kuwata and Martin, 2012; Shiraiwa et al., 2011; Zhang et al., 2018; Zhou et al., 2019). Global model calculations suggest that the phase state of atmospheric SOA may vary between liquid, semi-solid and solid in the planetary boundary layer, while SOA should be mostly in a glassy state in the free troposphere (Shiraiwa et al., 2017). The occurrence of glassy SOA in the free troposphere may promote ice nucleation and cloud droplet activation (Knopf et al., 2018; Slade et al., 2017) and facilitate long-range transport of toxic organic compounds contained in SOA (Mu et al., 2018; Shrivastava et al., 2017b).

The formation and properties of SOA are large sources of uncertainty in the current understanding of global air quality, climate change, and public health. The development of SOA models is among the most challenging problems in atmospheric chemistry (Tsigaridis et al., 2014). In most current air quality, atmospheric chemistry and climate models, the limiting step of SOA

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formation is assumed to be gas-phase oxidation of VOC to form semi-volatile and low volatile products. Thus, gas-phase oxidation is described kinetically, while gas-particle partitioning is often approximated by quasi-instantaneous equilibrium partitioning of the oxidation products (Pankow, 1994; Shrivastava et al., 2017a; Tsigaridis et al., 2014). The assumption of quasi-instantaneous gas-particle equilibration, however, is in question if particles are highly viscous, semi-solid or glassy - especially at low temperatures and low relative humidity (RH) (Li and Shiraiwa, 2019; Shiraiwa and Seinfeld, 2012). Experimental studies found kinetic limitations for gas uptake and particle evaporation at low RH (Liu et al., 2016; Perraud et al., 2012; Vaden et al., 2011; Yli-Juuti et al., 2017), but not for mixing in SOA at medium or high RH (Ye et al., 2016; Ye et al., 2018). An appropriate treatment of kinetic limitations depending on ambient conditions is critical for accurately reproducing particle size distribution dynamics in SOA growth (Shiraiwa et al., 2013a; Zaveri et al., 2018; Zaveri et al., 2020).

The dynamics of gas-particle partitioning have been considered in a wide range of atmospheric aerosol models, including aerosol dynamics models (Liu et al., 2019; McVay et al., 2014; Pandis et al., 1993; Riipinen et al., 2011; Zaveri et al., 2014), kinetic multilayer models (Berkemeier et al., 2016; Fowler et al., 2018; Roldin et al., 2014; Shiraiwa et al., 2012), GECKO-A (Aumont et al., 2005), the volatility basis set approach (Trump and Donahue, 2014; Trump et al., 2014), the statistical oxidation model (Cappa et al., 2016; Jathar et al., 2016), and particle evaporation models (Vaden et al., 2011; Yli-Juuti et al., 2017). Most model studies are using the Fuchs-Sutugin approximation of mass-transport kinetics at the gas-particle interface with a fixed mass accommodation coefficient that does not vary with particle phase state nor with the volatility and diffusivity of the investigated organic compounds. Molecular dynamics simulations (Julin et al., 2014; Von Domaros et al., 2020) and a recent SOA chamber study (Liu et al., 2019) suggest that the mass accommodation coefficients for semi-volatile organic molecules on organic substrates are close to unity. Measurement-derived mass accommodation coefficients reported from thermodenuder investigations of SOA volatility distributions, however, were one to three orders of magnitude lower (Kostenidou et al., 2018; Lee et al., 2010; Saleh et al., 2011).

Overall, the relations between particle phase state, mass accommodation, and the growth and atmospheric evolution of SOA have not yet been resolved and continue to be a subject of scientific debate. In this study, we investigate the influence of volatility, diffusivity, and particle phase state on the mass accommodation and gas-particle partitioning of organic compounds in

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SOA by detailed and simplified kinetic modeling approaches, comparing the Fuchs-Sutugin approximation to a detailed kinetic multilayer model (KM-GAP, Shiraiwa et al. 2012) as well as approximate and transient two-film model solutions (MOSAIC, Zaveri et al., 2014).

Theory and Methods

Traditionally, dynamic models of aerosol chemistry and physics describe the rate of gas-particle partitioning by a first-order gas-particle mass transfer coefficient (k_{gp} in s^{-1}) based on the Fuchs-Sutugin approximation of gas diffusion in the transition regime (Seinfeld and Pandis, 2016):

$$k_{gp} = 4\,\pi\,D_g\,r_p\,N_p\,\beta \quad (1)$$

$$\beta = \frac{0.75\,\alpha\,(1+Kn)}{Kn^2+Kn+0.283\,Kn\,\alpha+0.75\alpha} \quad (2)$$

where D_g ($cm^2\ s^{-1}$) is the gas phase diffusivity, r_p (cm) is the particle radius, N_p (cm^{-3}) is the particle number concentration, Kn is the Knudsen number, and α is the mass accommodation coefficient which represents the probability for a gas molecule colliding with the surface of the particles to enter the condensed phase. Kn is the ratio of the mean free path in the gas phase (λ), which can be calculated using the mean thermal velocity (ω), and the particle radius: $Kn = \lambda / r_p = 3\,D_g / (\omega\,r_p)$ (Pöschl et al., 2007). Note that k_{gp} is also often termed as condensation sink (CS). According to the absorptive partitioning theory under the assumption of ideal mixing (Pankow, 1994; Trump and Donahue, 2014), the rate of change of the gas- and particle-phase mass concentrations (C^g, C^p) of an organic compound in SOA partitioning can be expressed as:

$$\frac{dC^g}{dt} = -k_{gp}\left(C^g - \frac{C^g}{C_{OA}}C^0\right) \quad (3)$$

$$\frac{dC^P}{dt} = k_{gp}\left(C^g - \frac{C^P}{C_{OA}}C^0\right) - k_b C^P \quad (4)$$

where C_{OA} ($\mu g\ m^{-3}$) is the organic aerosol particle mass concentration, C^0 ($\mu g\ m^{-3}$) is the gas phase saturation mass concentration of the pure organic compound, and k_b (s^{-1}) is the first-order rate coefficient for its reaction in the particle bulk. The term $\frac{C^P}{C_{OA}}C^0$ represents gas-phase concentration of Z right at the surface and condensation is driven by gas-phase concentration gradient of Z between the gas and condensed phases.

While the term mass accommodation coefficient is widely used in atmospheric aerosol studies, its precise meaning is not always well defined. In particular, α is often applied without specifying if and how deep a molecule has to penetrate beneath the surface to have entered the

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condensed phase (adsorption vs. absorption). This aspect is usually not critical for liquid droplets with rapid surface-bulk exchange, fast bulk diffusion, and swift equilibration between the condensed phase and the surrounding gas phase. For viscous or solid particles, however, it can be essential to distinguish and resolve the kinetics of surface and bulk processes, including accommodation at the surface, transfer across the gas-particle interface, and further transport into the particle bulk (Kolb et al., 2010; Pöschl et al., 2007; Shiraiwa et al., 2012).

Building on the PRA kinetic model framework (Pöschl et al., 2007) and the kinetic multilayer model of surface chemistry and gas-particle interactions in aerosols and clouds (KM-GAP; Shiraiwa et al., 2012), we have derived an expression for the mass accommodation coefficient as a function of penetration depth into the particle bulk and related parameters (see step-by-step derivation in Appendix):

$$\alpha(x) = \alpha_s \frac{1}{1 + \frac{\alpha_s\, \omega\, C^0}{4\, D_b\, \rho_p} x \cdot 10^{-12} \frac{\mathrm{g}}{\mu\mathrm{g}} \frac{\mathrm{cm}^{-3}}{\mathrm{m}^{-3}}} \qquad (5)$$

Here ω (cm s^{-1}) is the mean thermal velocity of the organic compound in the gas phase, D_b (cm^2 s^{-1}) is its diffusivity in the condensed phase, ρ_p (g cm^{-3}) is the particle density, and x (cm) is the penetration depth. The scaling factor 10^{-12} (g cm^{-3})/(μg m^{-3}) allows for inserting C^0 in the commonly used units of μg m^{-3}; it can be omitted when C^0 is inserted in g cm^{-3} or when all quantities are inserted with standard SI units (cgs or mks system of units).

The surface accommodation coefficient α_s, which corresponds to $\alpha(0)$ with the penetration depth of 0, is the probability for a gas molecule Z colliding with the surface not to be immediately scattered back to the gas phase but to be accommodated at the surface for period longer than the duration of an elastic scattering process (Pöschl et al., 2007). Various equivalent, similar or closely related terms and parameters have been defined and used in the scientific literature include (Kolb et al., 2010; Pöschl et al., 2007): the condensation coefficient (Pruppacher and Klett, 1997), adsorption coefficient (Shi et al., 1999; Turco et al., 1989; Worsnop et al., 2002), sticking coefficient (Hanson, 1997), sticking probability (Clement et al., 1996; Garrett et al., 2006), trapping probability (Masel, 1996), adsorptive mass accommodation coefficient (Elliott et al., 1991), and thermal accommodation coefficient (Li et al., 2001; Worsnop et al., 2002).

When the penetration depth equals one or two molecular layers, i.e., once or twice the effective molecular length or diameter (δ), the corresponding penetration-depth-dependent mass

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accommodation coefficient is equivalent to the quasi-static surface accommodation coefficient (α_{ss}) or bulk accommodation coefficient (α_b), respectively, as defined in earlier kinetic multilayer model studies (Shiraiwa et al., 2012): $\alpha(\delta) = \alpha_{ss}$ and $\alpha(2\delta) = \alpha_b$. A recent study has compared this kinetic multilayer (KM) modeling approach with molecular dynamics (MD) simulations to calculate mass accommodation coefficients for a variety of semi-volatile compounds with different volatilities in squalene (Von Domaros et al., 2020). The penetration depth was assumed to be equal to the sum of half of the molecule's own length and half of the length of a squalene molecule. For the evaluation of uncertainties and sensitivities, the penetration depth was also varied from semi-volatile molecule's own length as a lower bound to the half-width of the nonuniform free energy region determined by the MD free energy profile as an upper bound. Within this range, the results of MD and KM simulations were in good agreement with each other, confirming the consistency and validity of the multilayer approach (Von Domaros et al., 2020).

Using the two-film theory of mass transfer between gas and particle phase, Zaveri et al. (2014) showed that the effects of a concentration gradient in the particle can be represented by a thin film adjacent to the surface with the following thickness or effective penetration depth for non-reactive partitioning and reactive uptake, respectively:

$$x_{\mathrm{eff}} = r_{\mathrm{p}} / 5 \qquad \text{(non-reactive partitioning)} \qquad (6)$$

$$x_{\mathrm{eff}} = r_{\mathrm{p}} \left(\frac{1 - Q}{q \coth q - 1} \right) \qquad \text{(reactive uptake)} \qquad (7)$$

where Q is the ratio of the average particle-phase concentration to the surface concentration at steady state and q is a dimensionless diffusion-reaction parameter (Seinfeld and Pandis, 2016):

$$Q = 3 \left(\frac{q \coth q - 1}{q^2} \right) \qquad (8)$$

$$q = r_{\mathrm{p}} \sqrt{\frac{k_{\mathrm{b}}}{D_{\mathrm{b}}}} \qquad (9)$$

Note that q_Z is the ratio of the particle radius to the so-called reacto-diffusive length, $(D_b/k_b)^{0.5}$, representing the characteristic depth to which a species can penetrate while reacting in the particle bulk (Pöschl et al., 2007; Worsnop et al., 2002).

By inserting x_{eff} in equation (5), we obtain an effective mass accommodation coefficient that accounts for the influence of penetration depth and its dependence on the diffusivity and reactivity of the investigated chemical species in the particle:

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$$\alpha_{eff} = \alpha(x_{eff}) \quad (10)$$

Results and Discussion

To investigate and demonstrate the relevance of the kinetics of mass accommodation and the applicability of α_{eff}, we simulate the temporal evolution of partitioning and equilibration of semi-volatile organic compounds (SVOC) with C^0 = 100 µg m^{-3} and D_g = 0.1 cm^2 s^{-1} interacting with non-volatile seed particles with a number concentration of 5000 cm^{-3}, an initial diameter of 200 nm, and a surface accommodation coefficient $\alpha_s = \alpha(0) = 1$. For the SVOC, we assume initial gas- and particle-phase concentrations of 2 µg m^{-3} and 0 µg m^{-3}, respectively. The particles are assumed to be either liquid with a bulk diffusion coefficient $D_b = 10^{-7}$ cm^2 s^{-1} or semisolid with $D_b = 10^{-15}$ cm^2 s^{-1}.

Model calculations were performed with the detailed kinetic multilayer model of gas-particle interactions (KM-GAP, Shiraiwa et al. 2012), with the Model for Simulating Aerosol Interactions and Chemistry (MOSAIC; Zaveri et al., 2014), and with an aerosol dynamic model using the simple Fuchs-Sutugin gas-phase diffusion model (F-S) with different values of α_m. Here, the KM-GAP results can be regarded as a benchmark, because the KM-GAP model explicitly resolves all relevant processes - including gas diffusion, reversible adsorption, surface-bulk exchange, bulk diffusion, and condensed-phase reactions - and has been successfully validated by against experimental data of both non-reactive partitioning and reactive gas uptake in a wide range of aerosol and surrogate systems (Berkemeier et al., 2017; Shiraiwa et al., 2012). The MOSAIC model yields approximate and transient solutions building on a less detailed representation of gas-particle interactions, which does not resolve reversible adsorption and surface-bulk exchange (Zaveri et al., 2014). In the F-S approximation, the kinetics of particle-phase mass transport are represented only by α_m as inserted into Eq. (2).

For liquid particles with fast surface-bulk exchange and bulk diffusion ($D_b = 10^{-7}$ cm^2 s^{-1}), $\alpha(x)$ remains close to $\alpha_s = \alpha(0) = 1$, and all models yield the same result of fast mass transfer from the gas to particle phase and equilibration within one second. For semi-solid particles with $D_b = 10^{-15}$ cm^2 s^{-1}, however, the temporal evolution of the SVOC gas-phase and particle-phase concentrations varies between different models and different values of α as shown in Figure 1 on logarithmic scales.

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220 According to KM-GAP (black line), the initial uptake of SVOC by the semisolid particle
221 phase is as fast as approximated by F-S with $\alpha = \alpha_{ss} = \alpha_Z(\delta) = 3\times10^{-2}$ corresponding to a
222 penetration depth of only one molecular length (monolayer) below the particle surface. After one
223 second, however, the KM-GAP uptake is limited by bulk diffusion and slows down substantially.
224 After about one hour, KM-GAP converges with the F-S approximation using $\alpha = \alpha_{eff} = \alpha(r_p/5) =$
225 8×10^{-4}. Notably, the F-S approximation with α_{eff} is identical to the MOSAIC approximation,
226 although the latter is based on different rate equations using a unity mass accommodation
227 coefficient like KM-GAP ($\alpha_s = 1$) and a two-film approach of bulk diffusion (Zaveri et al., 2014).
228 The MOSAIC transient solution exhibits a very high and likely overestimated initial uptake
229 corresponding to the F-S approximation with $\alpha = \alpha_s = 1$, because it does not resolve reversible
230 adsorption and desorption at the surface (Shiraiwa et al., 2012). After ~1 min, however, the
231 MOSAIC transient solution converges with KM-GAP. Overall, Figure 1 demonstrates that
232 accurate modeling of SVOC partitioning and uptake into semisolid particles requires an explicit
233 treatment of reversible adsorption and desorption at short time scales (< 1 min) and an explicit
234 treatment of bulk diffusion at intermediate time scales (~1 min to ~1 h). At long timescales (> 1
235 h), the partitioning is reasonable well captured by both the MOSAIC approximation using a two-
236 film approach of bulk diffusion (Zaveri et al., 2014) as well as the simple F-S approximation
237 accounting for the influence of penetration depth with the effective mass accommodation
238 coefficient, α_{eff}, newly introduced this study.

239 Figure 2a shows the temporal evolution of the gas-phase concentration of organic
240 compounds with different volatilities (C^0 = 0.1 to 1000 μg m^{-3}) that undergo non-reactive
241 partitioning into semisolid seed aerosol particles (D_b = 10^{-15} cm^2 s^{-1}). At short timescales,
242 substantial deviations can occur for semi-volatile compounds (C^0 = 1 to 100 μg m^{-3}), but at longer
243 time scales KM-GAP and the F-S approximation with α_{eff} are in reasonably good agreement
244 (relative deviations <10% after ~1 h). For low-volatile compounds (C^0 < 1 μg m^{-3}), equilibration
245 is achieved faster than for semi-volatile compounds because local thermodynamic equilibrium
246 between the gas phase and the particle surface is quickly established by condensation without
247 significant re-evaporation (Li and Shiraiwa, 2019; Zaveri et al., 2014). Semi-volatile compounds
248 with reactive functional groups can undergo particle-phase reactions such as dimerization and
249 oligomerization (Ziemann and Atkinson, 2012). Peroxide-containing highly oxidized molecules

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(HOM) are labile with chemical half-lives shorter than one hour (Krapf et al., 2016; Tong et al., 2016; Tong et al., 2019), and a recent study has shown that particle-phase reactions must be considered to describe HOM effects on particle growth (Pospisilova et al., 2020). Model results for SVOC partitioning plus reactive uptake with different rate coefficients in semisolid aerosol particles are shown in Figure 2b. The results of the Fuchs-Sutugin approximation with $\alpha_{eff} = \alpha(x_{eff})$ and x_{eff} from Eq. 7 are identical to the MOSAIC approximate and transient solutions. The uptake predicted by KM-GAP is similar but slightly slower in case of high bulk reaction rate coefficients, which can be attributed to the influence of reversible adsorption and desorption at the surface. Additional simulations with $\alpha_s = 0.1$ confirm that the results of the Fuchs-Sutugin approximation with α_{eff} and the MOSAIC approximate solution are identical, and that the results of KM-GAP and the MOSAIC transient solution are similar (Fig. S1).

For a given surface accommodation coefficient of $\alpha_s = 1$, which is likely a good approximation for SVOC on organic surfaces (Julin et al., 2014; Von Domaros et al., 2020), Figures 3a and 3b show how the effective mass accommodation coefficient α_{eff} depends on volatility and bulk diffusivity as related to particle phase state and viscosity according to the Stokes-Einstein relation (Shiraiwa et al., 2011). In the liquid phase with high bulk diffusivity ($D_b > 10^{-10}$ cm^2 s^{-1}), α_{eff} is essentially the same as α_s independent of volatility ($\alpha_{eff} \approx \alpha_s \approx 1$). With a decrease of bulk diffusivity in viscous or semisolid particles, α_{eff} decreases substantially for SVOC ($0.3 < C^0 < 300$ μg m^{-3}) and so-called intermediate volatility organic compounds (IVOC; $300 < C^0 < 3\times10^6$ μg m^{-3}) but not for LVOC ($3\times10^{-4} < C^0 < 0.3$ μg m^{-3}) and so-called extremely low-volatile organic compounds (ELVOC; $C^0 < 3\times10^{-4}$ μg m^{-3}). The reason why compounds with higher volatility are more strongly affected by particle phase state and diffusivity is that they are more likely to desorb back to the gas phase when diffusion into the bulk is slow. Compounds with lower volatility exhibit much lower desorption rates and are less likely to re-evaporate even if their diffusion into the bulk is slow. On the other hand, the influence of particle phase state and diffusivity increases with particle size because longer pathways of diffusion are required for effective accommodation, penetration, and absorption of gas molecules into larger particles as illustrated in Figures 3c and 3d.

The theoretically predicted influence of volatility on effective mass accommodation is consistent with a recent experimental study of α-pinene SOA reporting that the observed mass accommodation coefficients decreased from ~1 for low-volatile compounds to ~0.3 for semi-

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volatile compounds (Liu et al., 2019). Particle viscosity and bulk diffusivity were not reported for these experiments, but values around 10^{-13} to 10^{-14} cm^2 s^{-1} had previously been estimated for the diffusion coefficient of organic compounds in α-pinene SOA under dry conditions (Zhou et al., 2013). As illustrated in Figure 4a, theoretical predictions of α_{eff} assuming $\alpha_s = 1$ and $D_b = 10^{-12}$ to 10^{-14} cm^2 s^{-1} approximately capture the decrease and encompass the variability and uncertainty range of the experimentally derived mass accommodation coefficients reported by (Liu et al., 2019). Indeed, the observational α_m values reported in (Liu et al., 2019) and other experimental studies are usually obtained by fitting measurement data with the F-S approximation, and thus they should be directly compared to effective mass accommodation coefficient α_{eff} as derived by integration of the F-S approximation with detailed kinetic models of mass transport across the gas-particle interface. Figure 4b shows a wide range of other measurement-derived mass accommodation coefficients for various SOA and surrogate systems (data points/shaded areas) in comparison to generic values of α_{eff} (lines) calculated for characteristic experimental conditions ($\omega = 2.0\times10^4$ cm s^{-1}, $\rho_p = 1$ g cm^{-3}, and $r_p = 100$ nm, and $D_b = 10^{-19}$ to 10^{-5} cm^2 s^{-1}). As indicated by molecular dynamics simulations and related studies, the surface accommodation coefficient (adsorption probability) for semi-volatile or low-volatile organic compounds on organic surfaces is likely close to unity, $\alpha_s = 1$ (Julin et al., 2014; Von Domaros et al., 2020). Accordingly, low observational values of α can be attributed to the penetration-depth dependence of mass accommodation and plausibly explained by different scenarios/combinations/ratios of volatility and diffusivity, which can lead to a substantial decrease of α_{eff} relative to α_s in semi-solid particles. With regard to the dependence of α_{eff} on C^0, mixing effects and non-ideality may lead to deviations between C^0 and C^* (Zuend and Seinfeld, 2012), which should be taken into account in further investigations of mass accommodation and its influence on the formation and growth of SOA particles.

On the other hand, high reactivity can compensate the influence of low diffusivity and mass transport limitations in the particle phase, keeping α_{eff} close to α_s. In case of non-reactive partitioning, the effective penetration depth used to calculate α_{eff} is one fifth of the particle radius, i.e., $x_{eff}/r_p = 0.2$ (Eq. 6). In case of reactive uptake, however, x_{eff} decreases with increasing reactivity and with decreasing diffusivity according to Eqs. (7) to (9). Figure 5a illustrates how the effective penetration depth normalized by particle radius, x_{eff}/r_p, decreases with increasing first-

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order bulk reaction rate coefficient, k_b, and with decreasing diffusion coefficient, D_b. The reduced effective penetration depths at high k_b and low D_b reflect that reactive uptake by semisolid particles proceeds mainly through chemical reaction near the surface (Shiraiwa et al., 2013a). Figure 5b illustrates how α_{eff} depends on volatility and diffusivity for reactive uptake with $\alpha_s = 1$ and a first-order bulk reaction rate coefficient $k_b = 0.1$ s^{-1}. In comparison to Fig. 3b for non-reactive partitioning, Fig. 5b shows that particle phase reactivity leads to an extension of the volatility-diffusivity parameter space where $\alpha_{eff} \approx \alpha_s$ (red area): For semi-solid particles with low diffusivity, the parameter range of strong deviations between α_{eff} and α_s (yellow/green/blue area) is shifted towards higher volatility (lower right corner).

Summary and conclusions

Traditional SOA modeling approaches are often using the Fuchs-Sutugin approximation of mass-transport kinetics at the gas-particle interface in combination with mass accommodation coefficients that are not appropriately defined, leading to inconsistent results and conclusions. To overcome such deficiencies and difficulties, we have introduced an effective mass accommodation coefficient α_{eff} that depends on penetration depth and is a function of surface accommodation coefficient, volatility, bulk diffusivity, and particle-phase reaction rate coefficient. Application of α_{eff} in the traditional F-S approximation of SOA partitioning yields results that are consistent with detailed kinetic multilayer models (KM-GAP; Shiraiwa et al., 2012) and two-film models (MOSAIC, Zaveri et al., 2014)).

We suggest that α_{eff} and its dependence on penetration depth and related parameters should be applied and considered when the F-S approximation is used to investigate and simulate gas-particle interactions in viscous or semi-solid organic aerosols. The simple parameterization can be incorporated into regional and global models for a more realistic representation of SOA processes in the atmosphere, which seems particularly important with regard to the ubiquity of amorphous semi-solid or glassy particles predicted for the free troposphere as well as planetary boundary layer air at low relative humidity and low temperature (Maclean et al., 2017; Shiraiwa et al., 2017).

In the analysis and interpretation of SOA chamber and laboratory experiments, α_{eff} provides a simply way of accounting for the potential impact of volatility, diffusivity, and particle phase state on the kinetics of gas-particle partitioning for analysis and interpretation of chamber experiments. In particular, it may help to address and resolve apparent inconsistencies between the

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definitions and parameter values of mass accommodation coefficients that are derived from experimental data and from molecular dynamics simulations.

At short timescales, however, α_{eff} is not sufficient to properly describe the kinetics of gas-particle interactions with the F-S approximation. Such conditions require detailed kinetic model simulations with kinetic multilayer models or equivalent approaches explicitly resolving mass transport at the surface and in the bulk of the particle. The same applies for particles with layered structures such as surface crusts (solid/viscous surface layers) that may form upon chemical aging and can strongly impact the uptake of semi-volatile compounds and multiphase chemical processes in the particle phase (Pfrang et al., 2011; Vander Wall et al., 2018; Zhou et al., 2019). Moreover, mixed organic-inorganic particles often undergo liquid-liquid phase separation at moderate and high RH (Krieger et al., 2012; You et al., 2014; Zuend and Seinfeld, 2012), and liquid-liquid phase separation can also occur for purely organic particles (Song et al., 2017). The interplay of particle phase state and phase separation can further impact SOA partitioning (Shiraiwa et al., 2013b). In such complex particle morphologies with multiple phases, gradients and discontinuities of diffusivity may occur within the particle bulk and require more advanced modeling approaches of gas-particle interaction kinetics to be addressed in future studies.

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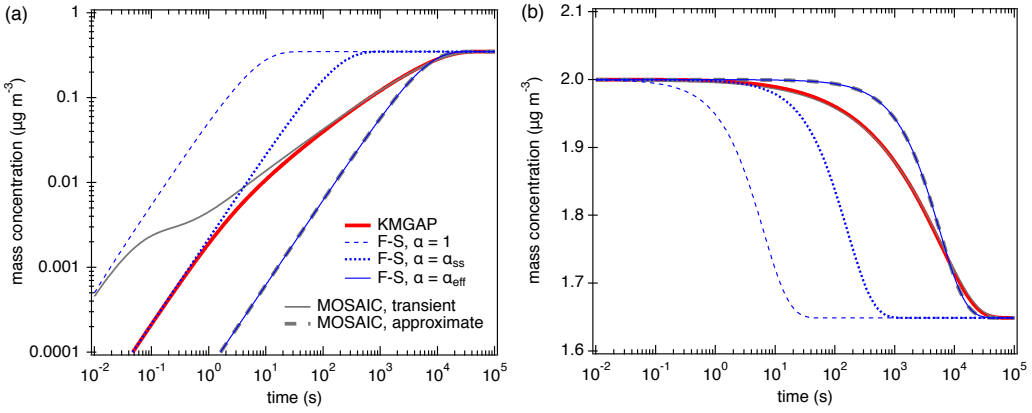

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360 **Figure 1.** Temporal evolution of the particle phase concentration (a) and the gas phase
361 concentration (b) of semi-volatile organic compounds (SVOC, C^0 = 100 μg m^{-3}) interacting with
362 semisolid seed aerosol particles (D_b = 10^{-15} cm^2 s^{-1}, ω = 2×10^4 cm s^{-1}, ρ_p = 1 g cm^{-3}). The red lines
363 are simulation results of KM-GAP, and the blue lines are the results of an aerosol dynamic model
364 that employs the Fuchs-Sutugin approximation with $\alpha = \alpha_s = 1$ (dashed), $\alpha = \alpha_{ss} = 3\times10^{-2}$ (dotted),
365 and $\alpha = \alpha_{eff} = 8\times10^{-4}$ (solid). The gray lines represent the MOSAIC transient solution (solid) and
366 approximate solution (dashed) (Zaveri et al., 2014).
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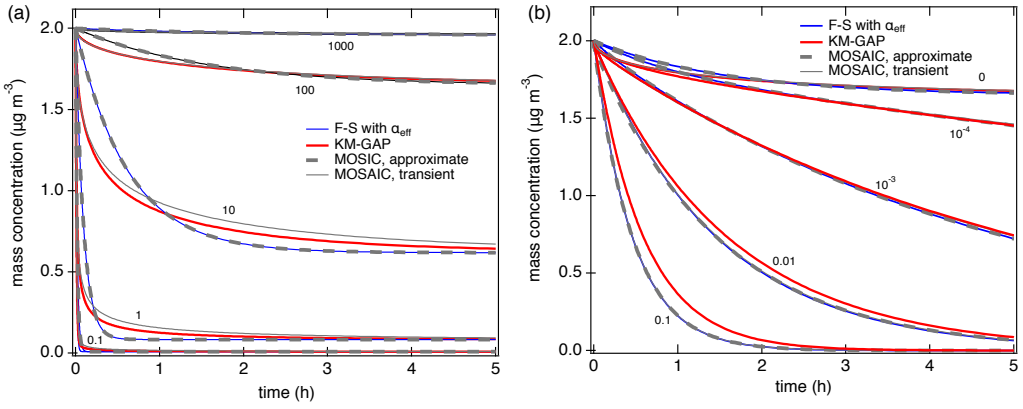

Figure 2. Temporal evolution of the gas phase concentration of organic compounds interacting with semisolid seed aerosol particles ($\alpha_s = 1$, $\omega = 2\times10^4$ cm s^{-1}, $D_b = 10^{-15}$ cm^2 s^{-1}, $\rho_p = 1$ g cm^{-3}): (a) Non-reactive partitioning of compounds with different volatilities (C^0 = 0.1, 1, 10, 100, 1000 μg m^{-3}); (b) reactive uptake of semi-volatile compounds (C^0 = 100 μg m^{-3}) with different first-order bulk reaction rate coefficients (k_b = 0, 10^{-4}, 10^{-3}, 0.01, 0.1 s^{-1}). The red lines are simulation results of KM-GAP, and the blue lines are the results of an aerosol dynamic model that employs the Fuchs-Sutugin approximation with $\alpha_{eff} = \alpha(r_p/5)$ for non-reactive partitioning (a) and with $\alpha_{eff} = \alpha(x_{eff})$ and x_{eff} from Eq. (5) for reactive uptake (b). The gray lines represent the MOSAIC transient solution (solid) and approximate solution (dashed) (Zaveri et al., 2014).

https://doi.org/10.5194/acp-2020-536
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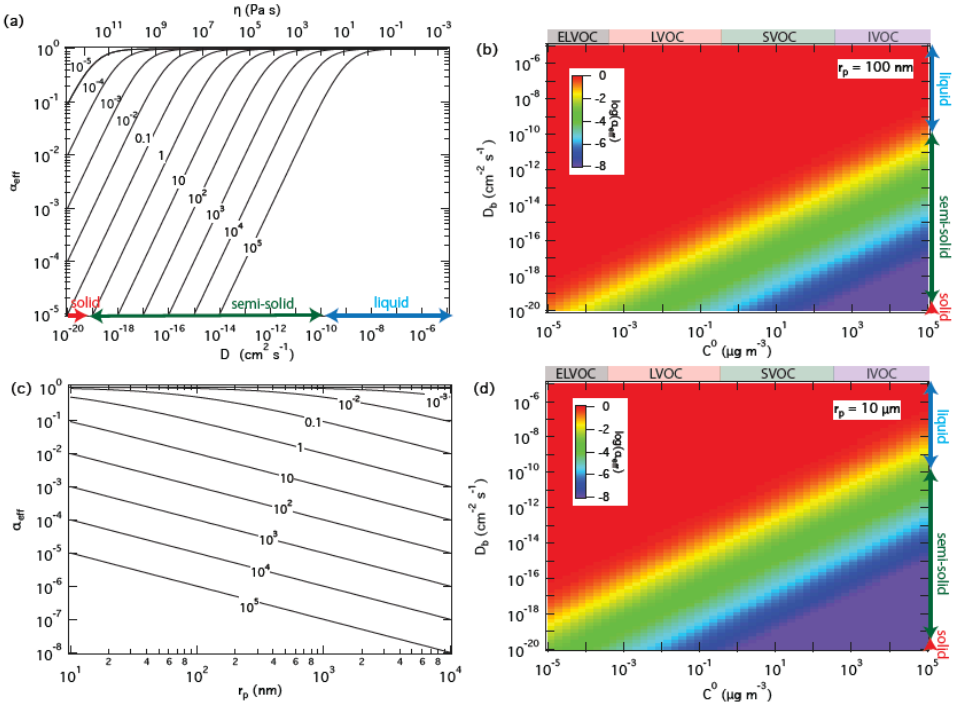

Figure 3. Effective mass accommodation coefficients, α_{eff}, for non-reactive partitioning of organic compounds Z ($\alpha_s = 1$, $\omega = 2\times10^4$ cm s^{-1}) with liquid, semi-solid, or solid aerosol particles ($\rho_p = 1$ g cm^{-3}) depending on pure compound volatility, C^0, particle bulk diffusivity, D_b (corresponding to viscosity, η), and particle radius, r_p: α_{eff} calculated as a function of D_b for $C^0 = 10^{-5}$ to 10^5 µg m^{-3} with $r_p = 100$ nm (a); α_{eff} calculated as a function of C^0 and D_b with $r_p = 100$ nm (b) and 10 µm (d); α_{eff} calculated as a function of particle radius for $D_b = 10^{-15}$ cm^2 s^{-1} and different levels of volatility ($C^0 = 10^{-3}$ to 10^5 µg m^{-3}) (c).

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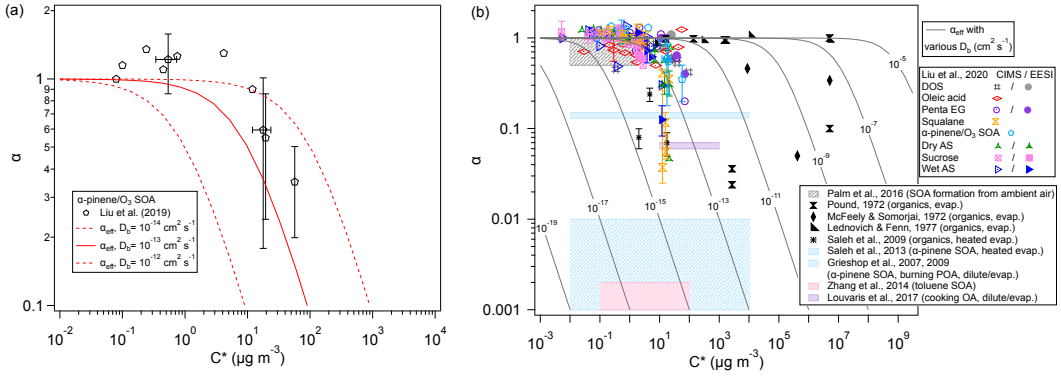

Figure 4. Effective mass accommodation coefficients, α_{eff} (lines, Eqs. 5-10) compared to measurement-derived mass accommodation coefficients, α (data points/shaded areas, Eqs. 1-2), plotted against effective saturation mass concentration, C^*, for various SOA and surrogate systems assuming $\alpha_s = 1$, $\omega = 2\times10^4$ cm s^{-1}, $\rho = 1$ g cm^{-3}, $r_p = 100$ nm, and $C^0 = C^*$: (a) observational results from laboratory experiments with semi-volatile components of SOA generated by ozonolysis of α-pinene (data points, (Liu et al., 2019)) compared to α_{eff} for $D_b = 10^{-14}$ to 10^{-12} cm^2 s^{-1} (lines); (b) observational results from earlier experimental investigations of laboratory-generated and ambient samples (data points/shaded areas, compiled by Liu et al., 2019) compared to generic values of α_{eff} for $D_b = 10^{-19}$ to 10^{-5} cm^2 s^{-1}(lines).

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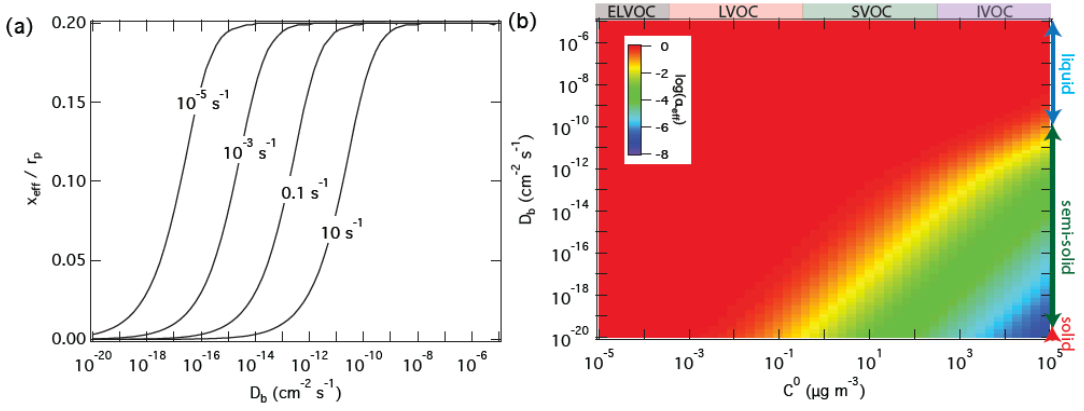

Figure 5. Effective penetration depths normalized by particle radius, x_{eff}, and mass accommodation coefficients, α_{eff}, for reactive uptake of organic compounds Z ($\alpha_s = 1$, $\omega = 2\times10^4$ cm s^{-1}) by liquid, semi-solid, or solid aerosol particles (r_p = 100 nm, ρ_p = 1 g cm^{-3}) depending on pure compound volatility, C^0, particle bulk diffusivity D_b, and first-order bulk reaction rate coefficient, k_b: (a) x_{eff} calculated as a function of D_b and k_b = 10^{-5} to 10 s^{-1}; (b) α_{eff} calculated as a function of C^0 and D_b for k_b = 0.1 s^{-1}.

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Appendix:

Derivation of penetration-depth-dependent mass accommodation coefficient

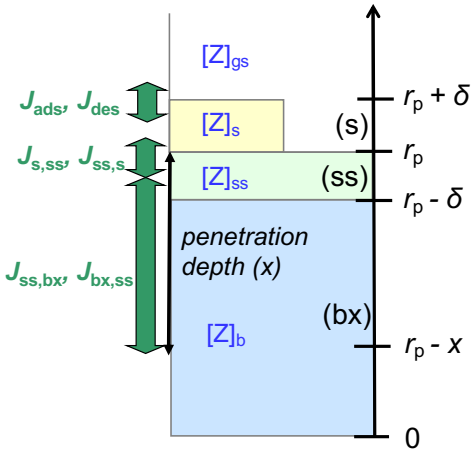

Figure A1. Schematic illustration of the kinetic multilayer modelling approach resolving mass transport fluxes (J) between the near-surface gas phase (gs), the sorption layer (s), the quasi-static surface layer (ss), and the bulk layer at penetration depth x (bx) (Shiraiwa et al., 2012).

Figure A1 illustrates the applied kinetic multi-layer model framework, in which the structure and composition of a particle are described by a sorption layer (s), a quasi-static surface layer (ss), multiple bulk layers (b), and any volatile, semi-volatile, or low-volatile chemical species (Z) that can undergo gas-particle partitioning and transport between the different layers and chemical reactions with each other (Pöschl et al., 2007; Shiraiwa et al., 2012). At low gas-phase concentration levels or high surface bulk exchange rates (e.g., for liquid particles under dilute atmospheric conditions), surface coverage and saturation effects can be neglected, and the surface accommodation coefficient (α_s) approaches the parameter value for an adsorbate-free surface ($\alpha_s \approx \alpha_{s,0}$) (Pöschl et al., 2007; Shiraiwa et al., 2012). In the absence of condensed-phase reactions, a quasi-static surface accommodation coefficient (α_{ss}), i.e. the probability for a gas molecule colliding with the surface to enter the quasi-static surface layer, can be calculated as follows (Shiraiwa et al., 2012):

$$\alpha_{ss} = \alpha_s \frac{J_{s,ss}}{J_d + J_{s,ss}} = \alpha_s \frac{k_{s,ss}}{k_d + k_{s,ss}} \quad \text{(A1)}$$

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Here J_d is the desorption flux of Z and k_d is the corresponding first-order rate coefficient; $J_{s,ss}$ and $k_{s,ss}$ represent the flux and first-order rate coefficient of transfer between the sorption layer and the quasi-static surface layer. The probability for an individual gas molecule colliding with the surface to enter the bulk with a penetration depth x can be described by a penetration depth-dependent mass accommodation coefficient, $\alpha(x)$, defined as follows:

$$\alpha(x) = \alpha_{ss} \frac{\Psi_{ss,bx}}{1-\Psi_{ss,s}\Psi_{s,ss}} \quad \text{(A2)}$$

Here $\Psi_{s,ss}$ is the probability for Z in the sorption layer to enter the quasi-static surface layer and $\Psi_{ss,bx}$ and $\Psi_{ss,s}$ are the probabilities for Z in the quasi-static surface layer to enter the bulk with the penetration depth of x or the sorption layer, respectively, which are determined by the corresponding fluxes and first-order rate coefficients of mass transport (Shiraiwa et al., 2012):

$$\Psi_{s,ss} = J_{s,ss}/(J_{s,ss} + J_{des}) = k_{s,ss}/(k_{s,ss} + k_d) \quad \text{(A3)}$$

$$\Psi_{ss,s} = J_{ss,s}/(J_{ss,bx} + J_{ss,s}) = k_{ss,s}/(k_{ss,bx} + k_{ss,s}) \quad \text{(A4)}$$

$$\Psi_{ss,bx} = J_{ss,bx}/(J_{ss,bx} + J_{ss,s}) = k_{ss,bx}/(k_{ss,bx} + k_{ss,s}) \quad \text{(A5)}$$

Inserting Eqs. (A3)-(A5) in Eq. (A2) leads to:

$$\alpha(x) = \alpha_s \frac{k_{s,ss}}{k_d + k_{s,ss}} \frac{\frac{k_{ss,bx}}{k_{ss,bx} + k_{ss,s}}}{1 - \frac{k_{ss,s}}{k_{ss,bx} + k_{ss,s}} \cdot \frac{k_{s,ss}}{k_{s,ss} + k_d}}$$

$$= \alpha_s \frac{k_{s,ss}k_{ss,bx}}{(k_d + k_{s,ss})(k_{ss,bx} + k_{ss,s})\left(1 - \frac{k_{ss,s}}{k_{ss,bx} + k_{ss,s}} \cdot \frac{k_{s,ss}}{k_{s,ss} + k_d}\right)}$$

$$= \alpha_s \frac{k_{s,ss}k_{ss,bx}}{(k_d + k_{s,ss})(k_{ss,bx} + k_{ss,s}) - k_{ss,s}k_{s,ss}}$$

$$= \alpha_s \frac{k_{s,ss}k_{ss,bx}}{k_d k_{ss,bx} + k_{s,ss}k_{ss,bx} + k_d k_{ss,s}} = \alpha_s \frac{1}{1 + \frac{k_d k_{ss,s} + k_d k_{ss,bx}}{k_{s,ss}k_{ss,bx}}}$$

$$= \alpha_s \frac{1}{1 + \frac{k_d}{k_{s,ss}} \frac{k_{ss,s} + k_{ss,bx}}{k_{ss,bx}}} = \alpha_s \frac{1}{1 + \frac{k_d}{k_{s,ss}}\left(1 + \frac{k_{ss,s}}{k_{ss,bx}}\right)} \quad \text{(A6)}$$

The first-order rate coefficients of adsorption and desorption are given by $k_a = \alpha_s \omega / 4$ and $k_d = 1/\tau_d$, respectively, where ω (cm s^{-1}) is the mean thermal velocity of Z in the gas phase and τ_d is the lifetime of desorption from the sorption layer (Pöschl et al., 2007; Shiraiwa et al., 2012). The rate

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456 coefficient of mass transfer between sorption layer and quasi-static surface layer can be estimated
457 based on the Fick's first law of diffusion considering that a molecule in the sorption layer needs to
458 travel a distance of δ to move into the quasi-static surface layer: $k_{ss,s} \approx D_b / \delta^2$ (Shiraiwa et al.,
459 2012). An estimate for $k_{s,ss}$ can be determined considering mass transport under equilibrium
460 conditions, where mass balance implies $J_{s,ss} = J_{ss,s}$, i.e., $k_{s,ss}\,[Z]_{s,eq} = k_{ss,s}\,[Z]_{ss,eq}$, and $J_{des} = J_{ads}$, i.e.,
461 $k_d\,[Z]_{s,eq} = k_a\,[Z]_{g,eq}$ (Shiraiwa et al., 2012). Here $[Z]_{g,eq}$, $[Z]_{s,eq}$, and $[Z]_{ss,eq}$ are the equilibrium or
462 solubility saturation number concentrations of Z in the gas phase, on the sorption layer, and in the
463 quasi-static surface layer, respectively:

464 $$k_{s,ss} = k_{ss,s} \frac{k_d\,[Z]_{ss,eq}}{k_a\,[Z]_{g,eq}} \quad \text{(A7)}$$

465 $$\frac{k_d}{k_{s,ss}} = \frac{k_a}{k_{ss,s}} \frac{[Z]_{g,eq}}{[Z]_{ss,eq}} = \frac{k_a}{k_{ss,s}} \frac{[Z]_{g,eq}}{[Z]_{b,eq}\,\delta} \quad \text{(A8)}$$

466

467 In analogy, the first-order rate coefficient $k_{bx,ss}$ can be estimated based on the Fick's first law of
468 diffusion, considering that a molecule Z at penetration depth x in the bulk needs to travel a distance
469 of $x - \delta$ to move into the quasi-static surface layer (Fig. A1): $k_{bx,ss} \approx D_b/(x - \delta)$. Under equilibrium
470 conditions, $J_{ss,bx} = J_{bx,ss}$ and $k_{ss,bx}\,[Z]_{ss,eq} = k_{bx,ss}\,[Z]_{b,eq}$ which leads to $k_{ss,bx} = k_{bx,ss}/\delta = D_b/(\delta\,(x - \delta))$
471 assuming ideal mixing conditions and $[Z]_{b,eq} = [Z]_{ss,eq}/\delta$ (Shiraiwa et al., 2012). Thus, $k_{ss,s} / k_{ss,bx} =$
472 $(D_b / \delta^2) / (D_b / (\delta\,(x - \delta))) = (x - \delta) / \delta$.

473 Based on the absorptive partitioning theory (Donahue et al., 2006; Pankow, 1994),

474

475 $$C^0 = \frac{C^g}{C^P} C_{OA} \quad \text{(A9)}$$

476 where C^0 (μg m^{-3}) is the pure compound saturation mass concentration, C^g and C^P (μg m^{-3}) are the
477 gas-phase and particle-phase mass concentrations of the compound Z, respectively, and C_{OA} (μg
478 m^{-3}) is the total organic aerosol mass concentration. C^g and $[Z]_{g,eq}$ are related through the following
479 equation:

480 $$C^g = \frac{[Z]_{g,eq} M}{N_A} \cdot 10^{12} \frac{\mu g}{g} \frac{m^{-3}}{cm^{-3}} \quad \text{(A10)}$$

481 where M is the molar mass of compound Z. $[Z]_{g,eq}$ is the equilibrium (saturation) number
482 concentration of Z in the gas phase. $[Z]_{g,eq}$ can be calculated using the saturation vapor pressure p:

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483 $[Z]_{g,eq} = p\ N_A\ /\ (R\ T)$ where N_A is the Avogadro number, R is the gas constant, and T is the
484 temperature. $[Z]_{b,eq}$ corresponds to the ratio between the number concentration of Z in the particle
485 phase (per m^3 of air) to the particle volume concentration (m^3 per m^3 of air), which can be
486 expressed using C_Z^{PM} and C_{OA} with the particle density ρ_P (g cm^{-3}):

487 $$[Z]_{b,eq} = \frac{\frac{C^P}{M} N_A}{\frac{C_{OA}}{\rho}} = \frac{C^P N_A \rho_P}{C_{OA} M} \quad (A11)$$

488 Combining Eq. (A9) – (A11) would lead to:

489 $$\frac{[Z]_{g,eq}}{[Z]_{b,eq}} = \frac{\frac{C^g N_A}{M}}{\frac{C^P N_A \rho_P}{C_{OA} M}} \cdot 10^{-12} = \frac{C^g}{C^P} C_{OA} \frac{1}{\rho_P} \cdot 10^{-12} = \frac{C^0}{\rho_P} \cdot 10^{-12} \frac{g}{\mu g} \frac{cm^{-3}}{m^{-3}} \quad (A12)$$

490 Inserting Eq. (A8) into Eq. (A6) and combination with Eq. (A12) leads to:

491 $$\alpha(x) = \alpha_s \frac{1}{1 + \frac{k_a}{k_{ss,s}} \frac{[Z]_{g,eq}}{[Z]_{b,eq}} \delta \left(1 + \frac{x - \delta}{\delta}\right)} = \alpha_s \frac{1}{1 + \frac{\alpha_s \omega C^0}{4 D_b \rho_P} x \cdot 10^{-12} \frac{g}{\mu g} \frac{cm^{-3}}{m^{-3}}} \quad (A13)$$

492

493 **Acknowledgements.** MS acknowledges funding by the National Science Foundation (AGS-
494 1654104) and the Department of Energy (DE-SC0018349). We thank Jose Jimenez (CU Boulder)
495 for stimulating discussions and for sharing published data and experimental information as
496 presented in Figure 4.

497

498 **Author contributions.** MS and UP designed the study, analyzed the data, and wrote the paper.
499 MS conducted kinetic modeling.

500

501 **Competing interests**. The authors declare that they have no conflict of interest.

502

503 **Data availability.** The simulation data may be obtained from the corresponding author upon
504 request.

505

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