



Heterogeneous nucleation of water vapor on different types of black carbon particles

Ari Laaksonen^{1,2}, Jussi Malila³, Athanasios Nenes^{4,5}

5

¹Finnish Meteorological Institute, 00101 Helsinki, Finland.

²Department of Applied Physics, University of Eastern Finland, 70211 Kuopio, Finland.

³Nano and Molecular Systems Research Unit, 90014 University of Oulu, Finland.

10 ⁴Laboratory of Atmospheric Processes and their Impacts, School of Architecture, Civil & Environmental Engineering, École Polytechnique Federale de Lausanne, 1015, Lausanne, Switzerland

⁵ Institute of Chemical Engineering Sciences, Foundation for Research and Technology Hellas (FORTH/ICE-HT), 26504, Patras, Greece

Correspondence to: Ari Laaksonen (ari.laaksonen@fmi.fi)

15 **Abstract:** Heterogeneous nucleation of water vapor on insoluble particles affects cloud formation, precipitation, the hydrological cycle and climate. Despite its importance, heterogeneous nucleation remains a poorly understood phenomenon that relies heavily on empirical information for its quantitative description. Here, we examine heterogeneous nucleation of water vapor on and cloud drop activation of different types of soots, both pure black carbon particles, and black carbon particles mixed with secondary organic matter. We show that the recently developed adsorption nucleation theory
20 quantitatively predicts the nucleation of water and droplet formation upon particles of the various soot types. A surprising consequence of this new understanding is that, with sufficient adsorption site density, soot particles can activate into cloud droplets – even when completely lacking any soluble material.



25 1 Introduction

When water vapor becomes supersaturated – i.e., its relative humidity exceeds 100% – it is in a metastable state, and can form liquid water or ice. External surfaces can facilitate the phase transition, through a process known as heterogeneous nucleation. Industrial processes (e.g. dropwise condensation), biological systems (e.g. infection strategy of plant pathogens), everyday situations (e.g. fogging of glasses and windshields), and fundamentally important atmospheric phenomena (clouds, frost and hoar frost) are all driven by the heterogeneous nucleation of water (Pruppacher and Klett, 1997; Franks, 2003).

Despite its ubiquity and importance, heterogeneous nucleation of water vapor remains poorly understood even after more than a century of research (Möller, 2008). This poor understanding is expressed by the lack of an established heterogeneous nucleation theory that provides quantitative comprehension of the process, and translates to a large uncertainty regarding the role of aerosol–cloud interactions in the climate system (Seinfeld et al., 2016). All traditional formulations are variants of classical nucleation theory (CNT) (Fletcher, 1958), which provide notoriously poor predictions of water drop nucleation (Mahata and Alofs, 1975). Molecular simulations (Lupi et al., 2014) are able to reveal aspects of heterogeneous nucleation phenomena, but on their own cannot provide a theoretical framework for describing heterogeneous nucleation in atmospheric and climate models – and are impractical for implementation in models themselves.

A shortcoming of CNT is that it does not recognize the existence of any water on surfaces prior to the formation of a macroscopic droplet (Fig. 1a). Another shortcoming is that all the energetics of interaction between the surface and the nucleating droplet are expressed by one parameter, the contact angle (Fletcher, 1958). In reality, water adsorbs to the surface already at sub-saturated conditions (relative humidity below 100%), which can play a critical role in the nucleation process. Notably, the energetics of adsorbed water differs considerably from what is implied by the single value provided by the contact angle. In fact, the heat of adsorption changes with increasing adsorption layer thickness (Hill, 1949a). Not describing the process of water adsorption therefore omits important physics that may hinder a quantitative description of heterogeneous nucleation. Unlike CNT, the adsorption nucleation theory (ANT) (Laaksonen, 2015; Laaksonen and Malila 2016) accounts for water adsorption prior to onset of heterogeneous nucleation (Figs. 1a and 1b).



Soot particles formed in combustion processes often contain sufficient amounts of water-soluble impurities so that their CCN activity can be described using the kappa-Köhler theory (Petters and Kreidenweis, 2007), provided that the kappa-value is known. An exception is the major fraction of soot particles produced by aircraft engines, which contain only small amounts of impurities, and are rather hydrophobic (Demirdjian et al., 2006). As those particles may participate in contrail or cirrus ice formation (Ikhenaze et al, 2020) either directly after emission or at some later point in time, it is important to be able to characterize their water uptake and nucleation properties theoretically as well as possible.

Water insoluble atmospheric particles often have rough surfaces characterized by cavities that display physical and chemical heterogeneity. This means that nucleation is not equally likely for all locations on the surface, which is another challenge for CNT. Here we show that when both adsorption and appropriate geometric considerations are included in the theoretical description of water nucleation, a quantitative theory of heterogeneous nucleation of water on aerosol is established. We demonstrate the new framework for different types of black carbon (BC) particles, which are long known to be climate active – but with a large uncertainty on the cumulative impact.

60

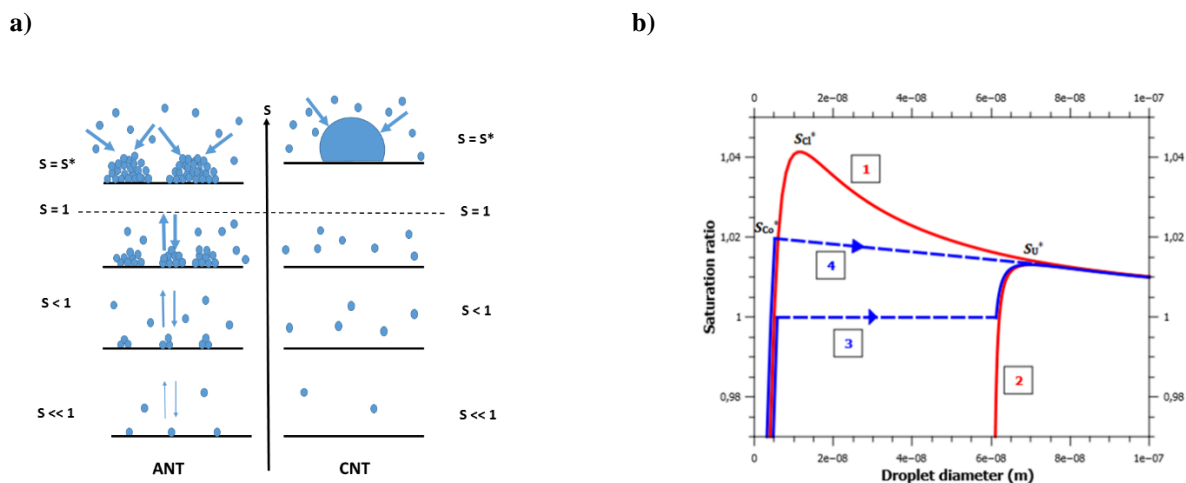


Figure 1. a) A schematic showing the principal differences between ANT and CNT. ANT is based on the premise that
65 vapor molecules are adsorbed on the surface already at subsaturation, and the molecular nature of adsorption is
reflected in the values of the FHH parameters A and B that are obtained from experimental adsorption equilibria. The
premise of CNT is that liquid droplets appear on the surface once the vapor becomes supersaturated. When
experimental contact angles are used. CNT predicts higher critical supersaturations S^* , at which unlimited
condensation starts, than ANT. In practice, the contact angle used for CNT is an adjustable parameter to fit
70 experiments, while for ANT it is the bulk (observed) value, adjusted owing to the interaction of adsorbed water
molecules on the surface. b) Schematic Köhler type curves for various heterogeneous nucleation pathways of water on
a 60 nm spherical particle. Curve 1 depicts the saturation ratio as a function of the diameter of individual droplets with
contact angle of 30° . Nucleation takes place at the maximum of the curve. Curve 2 is the Köhler curve for the FHH
activation mechanism, corresponding to a contact angle of 0° . Curve 3 corresponds to a situation where the droplets
75 fill the surface of the solid particle just below $S = 1$, and coalesce to form a uniform liquid film, which then grows along
curve 2 until nucleation occurs at the maximum. Curve 4 is otherwise similar to curve 3, but the droplets fill the surface
at $S = 1.02$ and nucleation occurs instantaneously because the saturation ratio is higher than the maximum of curve 2.
Note that the blue curves have been slightly offset along the x-axis for clarity.



2 Theoretical framework

80 It has been shown with atomic force microscopy (Cao et al., 2011) that adsorbed water exists as discrete “patches” (droplets) of liquid water, rather than a uniform film, on non-wettable surfaces. These patches then grow with increasing relative humidity (RH). Unless they fill the surface and coalesce to form a film, the growth continues past 100% RH to the supersaturated regime (Rowley and Innes, 1942), owing to the Kelvin effect, which causes the equilibrium vapor pressure to increase over curved surfaces. At a characteristic level of supersaturation (the so called “critical supersaturation”), however, 85 the droplets experience heterogeneous nucleation, where unconstrained condensation of water vapor ensues to form a water droplet.

To account for the above considerations, we develop a theory that is based on physisorption and captures the progressive accumulation of water, from monolayer “patches” on a surface (at low RH) up to macroscopic amounts of condensed material throughout the adsorption–nucleation transition (at the critical supersaturation). For this, we adopt a multilayer 90 formulation that has minimal coefficients and can be informed by standard adsorption experiments. Our formulation (Laaksonen, 2015; Laaksonen and Malila, 2015; Laaksonen et al., 2016) uses a combination of the Frenkel–Halsey–Hill (FHH) multilayer adsorption theory (Frenkel, 1946; Halsey, 1948; Hill, 1949b) and the Kelvin effect (Defay et al., 1966). We treat the surface as covered with individual, adsorbed droplets with an average distance s from their nearest neighbours. If the average droplet size and the average distance between them is known, the experimentally observable adsorption layer 95 thickness can be calculated. Surface heterogeneities imply droplets with a distribution of radii and contact angles; but as a first approximation we assume all droplets to have the same size and contact angle (these can be considered as average values). The equilibrium condition for a single droplet is given by

$$\ln(S) = -\frac{A}{N_d^B} + \frac{2\gamma v_w}{kTR} \quad (1)$$

where S denotes saturation ratio, γ is surface tension, v_w is the volume of a water molecule, k is the Boltzmann constant, T is 100 temperature, R is the droplet radius, and A and B are FHH parameters. The average number of monolayers, $\overline{N_d}$, over the droplet projection area on the adsorbent surface is given by $\overline{N_d} = \overline{\delta} / \delta_m$, where $\overline{\delta}$ denotes the average water layer



thickness, and δ_m is the thickness of a monolayer given as v_w/σ , where σ is the cross-sectional area of an adsorbed water molecule. $\overline{N_d}$ and R can be related via a relation that depends on the contact angle θ and the curvature of the adsorbent surface (see Materials and Methods for details of the theory). Thus, unlike in CNT, contact angle is a purely geometric
105 parameter describing the behavior of bulk water on the surface. Molecular simulations have shown that the contact angle of a water nanodroplet on graphitic surface equals the macroscopic value for droplets having radii larger than about 4 nm (Sergi et al, 2012). For nanodroplets on hydrophilic clays, the contact angle does not vary for droplets containing 256 – 1000 molecules (Zheng et al., 2017) – which means that any change in contact angle refers to nanodroplets with radii of the order of 1 nm. The sizes of critical clusters calculated using our theory are typically much larger than that, justifying the use of
110 experimental contact angles in the heterogeneous nucleation calculations. In case of a zero contact angle, water forms a uniform film on the adsorbent surface. If uniform adsorption occurs on spherical particles with radius R_p , the average number of monolayers is $\overline{N} = (R-R_p)/\delta_m$ (Sorjamaa and Laaksonen, 2007; Laaksonen and Malila, 2016).

The FHH interaction parameters A and B describe the strength of molecular interaction between the adsorbent and the first adsorption layer, and the exponential decay of the interaction strength as a function of distance, respectively. With the van
115 der Waals model fluid, A and B have well defined values (note, however, that A depends on the assumed molecular cross-section of the adsorbing molecule), but in real systems they must be treated as empirical constants that can be determined from measured adsorption isotherms. In case of droplet-wise adsorption, the macroscopically determined adsorption layer thickness is a statistical measure that can be related to the droplet size if the average distance between the adsorbed droplets, s , and the contact angle are known. In practice, s can be obtained together with A and B from the adsorption isotherms
120 (Laaksonen, 2015; Laaksonen and Malila, 2016).

We note that there is considerable uncertainty with the value of the experimental contact angle, due to contact angle hysteresis (Tadmor, 2004). The difference between the advancing and receding angles can be on the order of 10 degrees or even more. In any case, with graphite the critical supersaturation is surprisingly insensitive to an uncertainty of around $\pm 5^\circ$ in the value of θ . The reason for this is that the contact angle is included in the equations used when the theory is fitted to
125 adsorption experiments in order to determine the values of the FHH parameters A and B . Thus, if the contact angle is



changed by a few degrees, the model has to be re-fitted to the adsorption data, which can be done simply by changing the value of the A -parameter. It turns out that this does not either compromise the quality of the fit to the adsorption experiments or noticeably change the calculated critical supersaturation unless the contact angle is changed by more than about plus or minus five degrees (see Fig. A2).

- 130 Figure 1b shows a schematic picture of nucleation of water vapor on a spherical particle in our model framework. A droplet grows along the red curve 1 until it reaches the critical supersaturation S_{cl}^* and nucleates. However, if the average distance between the droplets is small enough, and they coalesce to form a uniform liquid film at sufficiently low saturation ratio (blue curve 3), the new equilibrium as well as the new critical supersaturation S_U^* can be calculated simply by using a zero contact angle (red curve 2). Depending on the value of s , it is also possible that the droplet coalescence occurs at some
- 135 supersaturation S_{co}^* between S_U^* and S_{cl}^* , in which case nucleation is immediate (blue curve 4).

2.1 Adsorption nucleation theory

- The starting point of the theory is that adsorbing vapor molecules are assumed to cluster around adsorption sites to form spherical caps of liquid drops having a contact angle θ . The equilibrium condition is given by Eq. (1). The geometry for a
- 140 spherical seed particle is as shown in Fig. 2. The average over δ^B is given by

$$\overline{\delta^B} = [1 - \cos \Phi]^{-1} \int_0^\Phi \delta^B \sin \alpha d\alpha \quad (2)$$

with

$$\delta = -R_p + R \cos \beta + \sqrt{(R \cos \beta)^2 - R^2 + d^2} \quad (3)$$

$$d = \sqrt{R_p^2 + R^2 - 2R_p R \cos \theta} \quad (4)$$

145 $\cos \Phi = (R_p - R \cos \theta) / d \quad (5)$

$$\cos \beta = (R_p + \delta - d \cos \alpha) / R \quad (6)$$



S can now be computed as a function of R using Eq. (1) with $\overline{N_d^B} = \overline{\delta^B} / \delta_m^B$ evaluated numerically from Eqs. (2) – (6). The resulting curve shows a maximum at the critical supersaturation S^* (and critical radius R^*), marking the onset of heterogeneous nucleation.

150

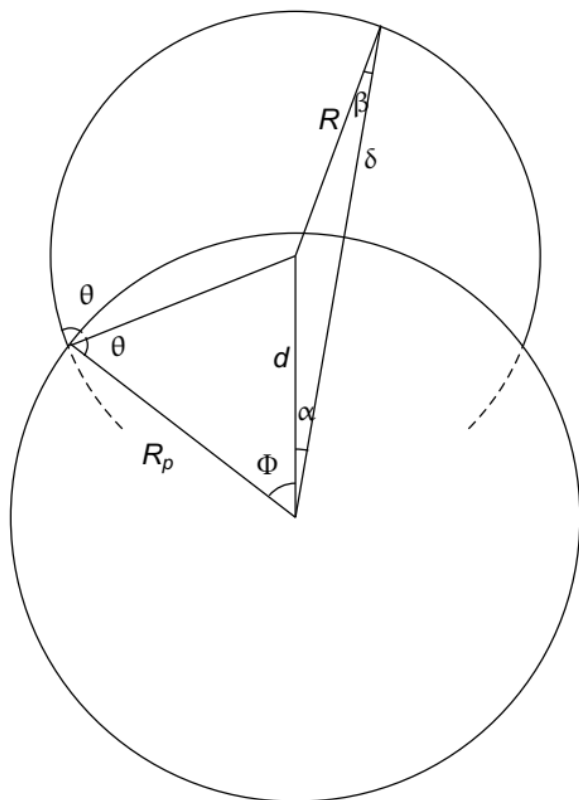


Figure 2. The geometry used in deriving equations for nucleation of a spherical drop on a spherical nanoparticle.

The values of the adsorption parameters A and B , as well as that of the distance between active sites s , can be obtained from
155 measured adsorption isotherms. If a portion of the adsorption data is clearly in the multilayer regime and the data points align linearly when $\ln(-\ln(S))$ is plotted vs. $\ln(N)$, parameters A and B can be determined by fitting the classical FHH equation $\ln(S) = -AN^B$ to the data (here, N denotes the macroscopically observable surface coverage). If the contact angle is known, the distance between active sites can be obtained from the sub-monolayer part of the adsorption isotherm. To do that, a relation is



needed between the size of an individual adsorbed droplet, and the macroscopic surface coverage. Making the approximation

160 $\overline{N_d^B} \approx \overline{N_d}^B$, Eq. (1) can be translated into

$$\ln(S) = -A \left[\frac{\pi \zeta^2}{\eta^2 N} \right]^{B/3} + \frac{2\gamma}{3kT} \left[\frac{\pi \eta \zeta^2}{N} \right]^{1/3}, \quad \theta \leq 90^\circ, \quad (7)$$

$$\ln(S) = -A \left(\frac{\eta}{\beta} \right)^B \left[\frac{\pi g(\theta)}{3s^2 N} \right]^{B/3} + \frac{2\gamma v}{3kT} \left[\frac{\pi g(\theta)}{3s^2 N} \right]^{1/3}, \quad \theta > 90^\circ \quad (8)$$

with $\zeta = 3v \sin \theta / s$, $\eta = \sigma f(\theta)$, $g(\theta) = 4 - (1 + \cos \theta)^2 (2 - \cos \theta)$, and $f(\theta) = \frac{(1 - \cos \theta)^2 (2 + \cos \theta)}{\sin^2 \theta}$, $\theta \leq 90^\circ$;

$$f(\theta) = 2 - 3 \cos \theta, \quad \theta > 90^\circ.$$

165 Distance between adsorption sites s can now be determined by fitting Eq. (7) or (8) to the sub-monolayer data (note that it is best to use the portion of the data just below a monolayer, as the low coverage data may be affected by pores or other factors not taken into account in the above equations.

When the adsorbent surface is poorly wettable, the adsorption data may not extend to the multilayer regime at all. In such a case, all of the parameters need to be obtained from a best fit of either Eq. (7) or Eq. (8) to the data set.

170 As noted in relation to Fig. 1b, nucleation can also occur due to coalescence of adsorbed droplets. The criteria we use for the coalescence nucleation are that (i) the projection areas of adsorbed droplets equal the surface area of the seed particle, and that (ii) the supersaturation exceeds the critical supersaturation calculated for a seed particle that is perfectly wettable.

2.2 Liquid drop nucleation on insoluble particles mixed with soluble material

175 If $\theta = 0^\circ$, i.e. the seed particle is completely wettable, the saturation ratio can be written in terms of dry and wet particle radii as

$$\ln(S) = -A \left(\frac{R - R_p}{D_w} \right)^{-B} + \frac{2\gamma v_w}{kTR} \quad (9)$$

where we have approximated monolayer thickness with the diameter of a water molecule, D_w (Laaksonen and Malila, 2016).

For a partially soluble seed particle – for example, a black carbon particle coated with low-volatility organic or inorganic



180 substances after atmospheric processing – Eq. (9) needs to be further modified to take the hygroscopicity of solution into
account. The following equation (Kumar et al, 2011) can be written in terms of the hygroscopicity parameter κ (Petters and
Kreidenweis, 2007):

$$\ln(S) = -A \left(\frac{R-R_p}{D_w} \right)^{-B} - \frac{\varepsilon_s R_p^3 \kappa}{(R^3 - \varepsilon_i R_p^3)} + \frac{2\gamma v_w}{kTR} \quad (10)$$

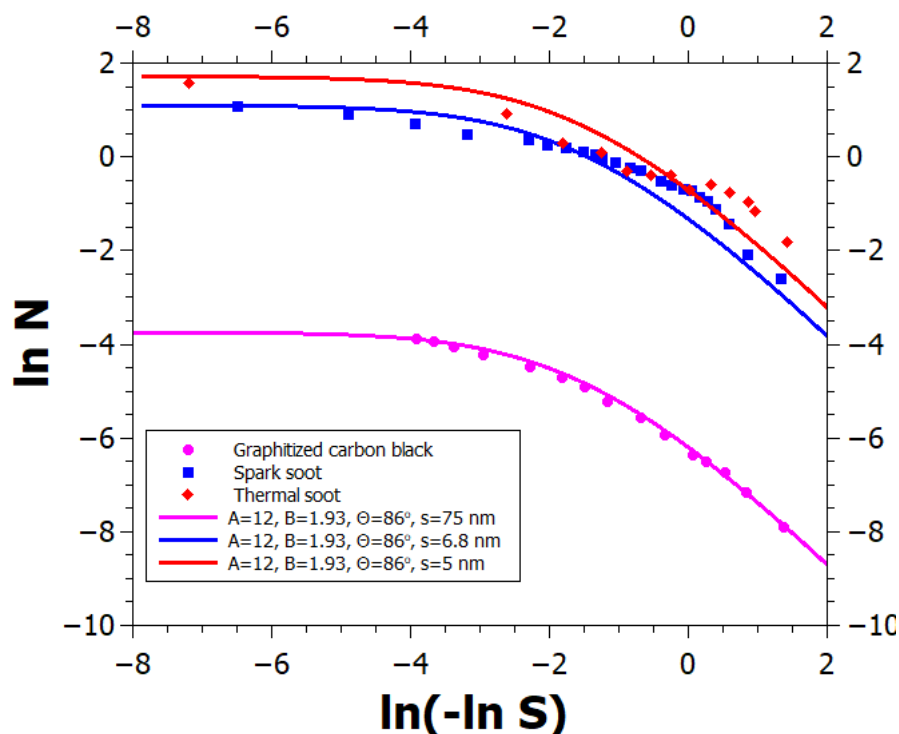
where ε_s and ε_i are the volume fractions of soluble and insoluble species in the seed particles.

185

3. Results and discussion

3.1 Constraining water adsorption parameters on carbon

We now focus on determination of water vapor adsorption parameters on graphite and on black carbon (BC) particles. Figure
2 shows a logarithmic FHH plot of several adsorption isotherms. Starting from the bottom, the lowermost measurement data
190 set (Young et al, 1954) shows adsorption on graphitized carbon black treated at 400 °C for 12 hours prior to the
measurements. Based on the minute amounts of water adsorbed at even the highest relative humidity, and information
obtained from heat of immersion measurements, Young et al. (1954) suggested that clusters of water adsorb at hydrophilic
surface sites which are far apart and may arise from surface oxides formed at the edge atoms of graphite, which has since
then been corroborated by quantum chemical calculations (Oubal et al., 2010). This description is in complete agreement
195 with the droplet-wise FHH adsorption model (see the lowermost solid line in Fig. 3 where the model reproduces the data).
The model parameters are shown in Fig. 3; the contact angle value was obtained from Fowkes and Harkins (1940), and the
parameters A , B , and s were determined by optimizing the model fit to the experimental data.



200 **Figure 3. An FHH-plot of experimental (markers) and modelled (lines) adsorption isotherms of water vapor on graphite and different soots. Note that in the abscissa, increasing supersaturation is to the left. See text for details.**

When BC is exposed to atmospheric conditions, the number of hydrophilic surface sites is expected to increase with time as different contaminants adsorb on the carbon and may undergo oxidation (Aria et al., 2016). Contaminants may also be present due to imperfect combustion, in case the carbon was produced by burning. This suggests that adsorption isotherms of BC containing impurities can be accounted for by increasing the density of the adsorption sites in the FHH model (without changing the other adsorption parameters), constrained by observations.

The data in the upper part of Fig. 3 (Kuznetsov et al., 2003; Popovicheva, 2008) show water adsorption on spark discharge soot, and thermal soot produced by natural gas pyrolysis. The materials were not treated at high temperature prior to the

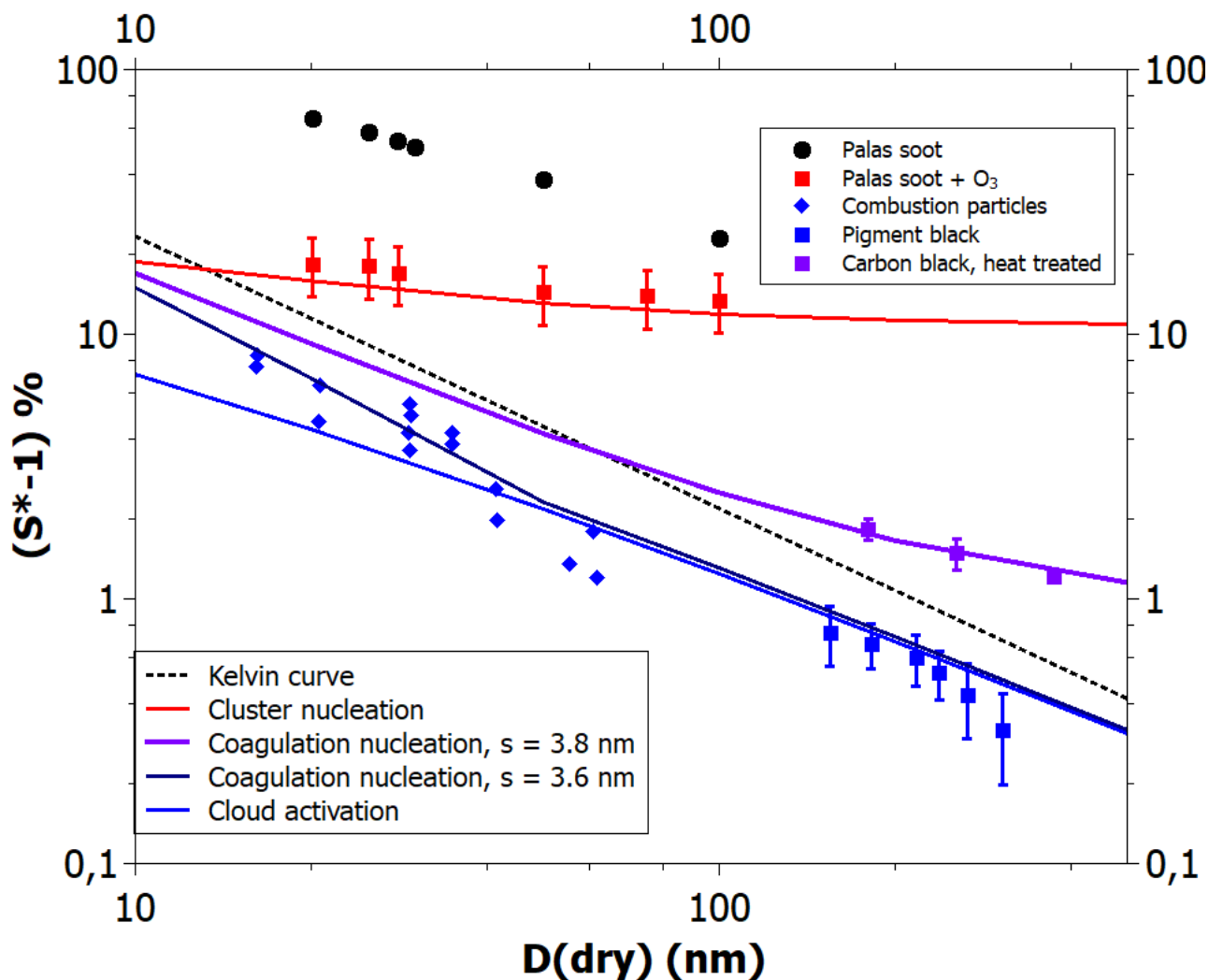


210 adsorption measurements, and are thus likely to contain more adsorption sites from surface contaminants than the
graphitized carbon black used by Young et al. (1954). The blue and red solid curves represent the corresponding adsorption
isotherms, obtained by simply reducing the average distance of adsorption sites in the FHH model from 75 nm to 6.8 nm and
5 nm, respectively, and keeping the other parameters at the graphitized carbon black values. The data for $RH > 80\%$ ($\ln(-\ln$
 $(S)) < -1.5$) agree well with the theoretical curves, whereas at lower relative humidities the data show more adsorption than
215 the model. The latter is caused by the effects of soot porosity, which is indicated by a hysteresis between the measured
adsorption and desorption curves.

When the density of hydrophilic adsorption sites is sufficiently high, the growing water clusters will start coalescing as the
RH increases above some threshold value, and a liquid film will be formed on the BC surface. If the film formation occurs at
RH below 100%, the conditions for activation of a BC particle into a cloud droplet (where the particle is said to act as a
220 Cloud Condensation Nuclei, or CCN) can be described using a zero contact angle and the total surface area of the particle,
while the full adsorption nucleation theory with non-zero contact angle is required when the active site density is low (curve
3 in Figure 1b). We next examine whether the CCN activation of different types of BC particles can be described with the
theoretical framework as constrained above, using model parameters shown in Fig. 3.

225 3.2 Water nucleation on black carbon particles

Figure 4 depicts CCN activation data for hydrophobic and hydrophilic BC particles. The most hydrophobic particles
(Kotzick et al., 1997) were produced using a spark discharger. The particles were in contact only with argon and nitrogen
gases prior to the activation measurements (black spheres), and thus contained very little adsorption sites. As can be seen, all
model curves are below the spark discharger data. However, the model with parameters obtained using the adsorption data
230 for graphitized carbon black (Fig. 3) is in excellent agreement with the CCN activation data for spark discharger particles
that were subjected to ozonolysis prior to the activation measurement (Kotzick et al., 1997) (red squares), and therefore have
an increased density of adsorption sites.



235 **Figure 4. Experimental (markers) and theoretical (lines) critical supersaturations as a function of particle size for cloud drop activation of hydrophobic and hydrophilic soot particles. See text for details.**

The blue diamonds in Fig. 4 show activation of soot particles produced by burning camp stove fuel (white gas) (Hagen et al., 1989). Such particles can be expected to contain a large amount of oxidized species on their surfaces, and may therefore already have a liquid film of water at activation. The blue curve shows model prediction with the same FHH parameters A and B as the red curve, but a zero contact angle. It appears that the model has somewhat gentler size dependence than the

240



rather scattered data, but otherwise the agreement is quite satisfactory. The CCN activation of pigment black particles (Dalirian et al., 2018), also match the theoretical line within uncertainty limits (with the exception of the largest particle size, which has the uncertainty range just below the theoretical line).

Very interestingly, the three datapoints of Dusek et al. (2006) are clearly above those of Dalirian et al. (2018). The principal
245 difference between the soots used in these two studies is that Dusek et al. (2018) applied heat treatment, thus very likely
reducing the adsorption site density. Indeed, a theoretical line calculated with average adsorption site distance 2.8 nm and
assuming the coagulation nucleation to occur, matches the data perfectly. When a site distance of just 0.2 nm smaller is
assumed, the theoretical prediction becomes equal to the CCN activation prediction for particles larger than about 50 nm.
With smaller particles, the coagulation nucleation appears to match the camp stove gas soot data better than the CCN
250 activation theory. In general, accounting for droplet-wise adsorption and resulting alternative pathways for heterogeneous
nucleation (Fig. 1) allows us to obtain a much-improved description on the nucleation of water on BC particles than CNT
(Kotzick et al., 1997) or adsorption nucleation theories not accounting for initial adsorption on distinct sites (Henson, 2007).
Further enhancement in BC hydrophilicity, on top of increased adsorption site density due to contaminant oxidation, may
occur e.g. as a result of heterogeneous chemistry leading to a substantial amount of water-soluble material on the BC
255 particles. The red diamonds in Fig. 5 (Wittbom et al., 2014) are for particles aged in a smog chamber prior to the activation
measurements, and were estimated to contain between 8 and 91 percent by mass of water soluble substance, respectively (the
data points were deciphered from Figs. 10 and 11 in Wittbom et al. (2014)). The blue crosses were calculated using the
theory which describes the combined effect of adsorption and water-soluble material when the volume fraction ε and the
hygroscopicity parameter κ of the water-soluble material are known (Kumar et al., 2011; Petters and Kreidenweis, 2007; see
260 previous section for details). The FHH parameters were kept the same as for the pure black carbon particles, and the κ -value
was assumed to be 0.13 (Wittbom et al., 2014). The model predictions, shown by the blue crosses, are once again in
excellent agreement with the data. Only in three cases the theory provides clear underestimates, the worst occurring at the
lowest SOA mass fraction of about 0.08, as can be seen in Fig. 5b, while Fig. 5a shows that the experimental point is very
close to the Kelvin line. Comparison to Fig. 4 indicates that this data point is compatible with activation of particles that do
265 not contain water soluble material.

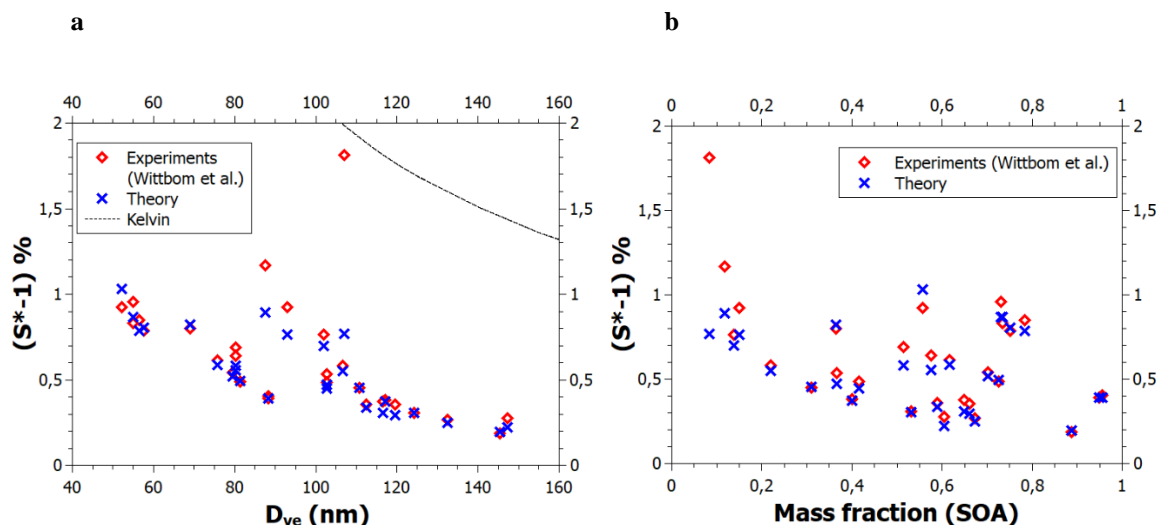


Figure 5. Experimental and theoretical critical supersaturations of water vapor on mixed soot particles containing BC and SOA as a function of a) volume equivalent diameter of dry particles and b) SOA mass fraction. The data are a subset of the data shown in Figures 10 and 11 of Wittbom et al. (2014), and were identified by matching supersaturations (to a precision $< 0.01\%$) in the two figures.

4. Conclusions

In this study, we have shown that the adsorption nucleation theory can predict the critical supersaturations for heterogeneous nucleation and CCN activation of water vapor on different types of black carbon particles, using a single set of experimental adsorption and contact angle parameters.

It is quite obvious that surface heterogeneities cause adsorbed droplets to have a range of radii and contact angles, but to make the theory usable, we assume that we can resort to average values. This applies to the FHH parameters A and B as well; it is likely that there would be variability if isolated surface regions were examined. However, the variability of the heterogeneities – whether chemical or physical – is reflected in the values of the “average” adsorption parameters obtained from macroscopic adsorption isotherms. Our hypothesis is that the same parameters can be made use of when extending the theory to the supersaturated regime to predict critical supersaturations. The good match between our theoretical predictions



and the experimental critical supersaturations provides clear indication that our hypothesis is well supported. This also implies that the heterogeneities occur at a spatial scale smaller than the surface area covered by the critical nuclei.

Our findings concerning adsorption nucleation on BC have a direct relevance for climate studies, as even the sign of the radiative forcing due to BC aerosol–cloud interactions remains uncertain (Bond et al., 2013). Notably, atmospheric lifetime of BC particles is regulated by condensation of water and subsequent cloud processes, which contributes further to the uncertainty of the total radiative forcing (Samseth et al., 2018).

Appendix A: Water properties used in the calculations

T denotes absolute temperature (K), T' is temperature in degrees Celsius, and R_g is the molar gas constant.

Equilibrium vapor pressure of liquid water (Pa) (Murphy and Koop, 2010):

$$p_w = \exp\{54.842763 - 6763.22/T - 4.21\ln(T) + 0.000367T + \tanh[0.0415(T-218.8)][53.878-1331.22/T-9.44523\ln(T)+0.014025T]\}$$

Density of liquid water (g cm^{-3}) (Marcolli, 2017):

$$\rho_w = 1.8643535 - 0.0725821489T + 2.5194368 \cdot 10^{-3}T^2 - 4.9000203 \cdot 10^{-5}T^3 + 5.860253 \cdot 10^{-7}T^4 - 4.5055151 \cdot 10^{-9}T^5 + 2.2616353 \cdot 10^{-11}T^6 - 7.3484974 \cdot 10^{-14}T^7 + 1.4862784 \cdot 10^{-16}T^8 - 1.6984748 \cdot 10^{-19}T^9 + 8.3699379 \cdot 10^{-23}T^{10}$$

Surface tension of water (N m^{-1}) (Hrubý et al., 2014):

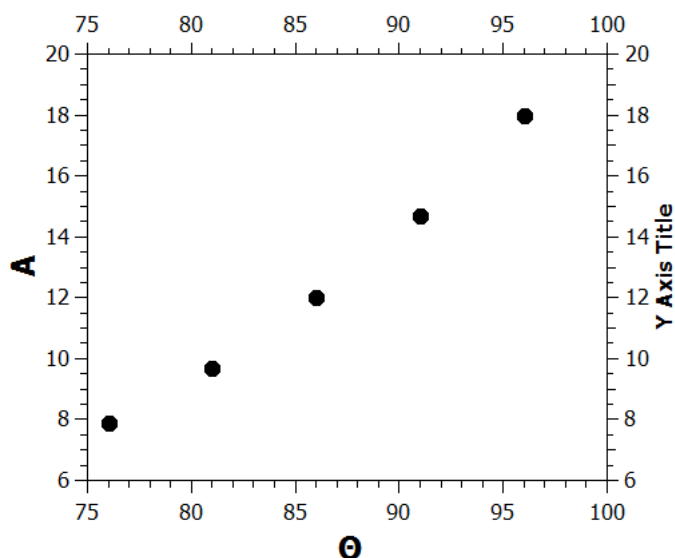
$$\gamma_w = 0.2358(1 - T/647.096)^{1.256}[1 - 0.625(1 - T/647.096)]$$

Appendix B: Sensitivity analysis

The contact angle value used in the calculations shown above is 86° ; here we examine how much the critical supersaturation is affected by changing the contact angle by up to plus/minus 10 degrees. As noted in the main text, when the contact angle is given a new value, the model needs to be re-fitted to the adsorption data by changing the value of the FHH parameter A . Figure A1 shows the values of A as a function of the contact angle in the range $76^\circ - 96^\circ$ when this is done (the values of the parameters B and s were kept at 1.93 and 75 nm, respectively). Figure A2 shows the experimental adsorption data together



with the model for $(\theta, A) = (76^\circ, 7.9)$, $(86^\circ, 12)$ and $(96^\circ, 18)$. As can be seen, all curves fit the data well, although there is some deviation at low saturation ratios. Figure S3 shows the calculated critical supersaturations for water nucleation. The dry particle diameter (100 nm) was chosen based on the experimental nucleation data of Kotzick et al. (1997). The relative uncertainty of the critical supersaturation of water due to an uncertainty of $\pm 5^\circ$ in θ is quite insignificant, about 1.5%.



310

Figure A1. Pairs of contact angle and the FHH parameter A that produce good model fits to the adsorption data.

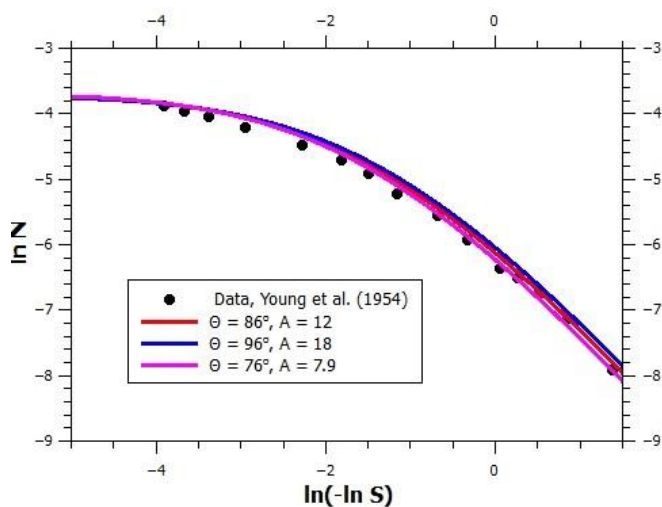
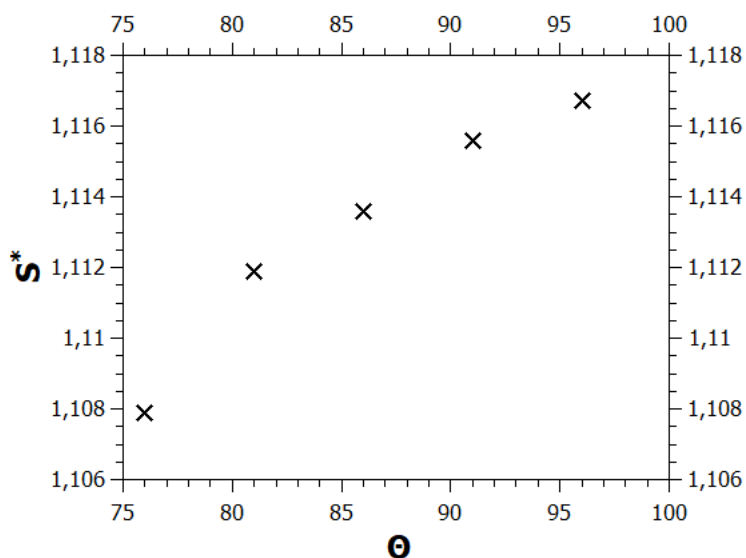


Figure A2. Model fits to adsorption data for three pairs of θ and A .

315



320 **Figure A3.** Critical supersaturations in water nucleation as a function of contact angle. The A -parameter values used in the calculations are as shown in Fig. S1. Nucleation temperature is +25 °C and dry particle diameter 100 nm.

Acknowledgments. This work was supported by the Academy of Finland, C-Main project (grant no. 309141) and the Center of Excellence programme (grant no. 307331), and by the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation programme (grant agreement No. 717022).

325 **Code and Data availability.** Code and model results are available upon request

Competing interests. The authors declare that they have no conflict of interest.

Author contributions. AL performed all theoretical calculations and drafted the manuscript. All co-authors discussed the results. JM and AN commented on the manuscript and contributed to writing of the final version.

References

330 Aria, A. I., Kidambi, P. R., Weatherup, R. S., Xiao, L., Williams, J. A., and Hofmann, S.: Time evolution of the wettability of supported graphene under ambient air exposure, *J. Phys. Chem. C* 120, 2215–2224, 2016



- Bond, T. C., Doherty, S. J., Fahey, D. W., Forster, P. M., Berntsen, T., DeAngelo, B. J., Flanner, M. G., Ghan, S., Kärcher, B., Koch, D., Kinne, S., Kondo, Y., Quinn, P. K., Sarofim, M. C., Schultz, M., Venkataraman, C., Zhang, H., Zhang, S., Bellouin, N., Guttikunda, S. K., Hopke, P. K., Jacobson, M. Z., Kaiser, J. W., Klimont, Z.,
335 Lohmann, U., Schwarz, J. P., Shindell, D., Storelvmo, T., Warren, S. G., and Zender, C. S.: Bounding the role of black carbon in the climate system: A scientific assessment, *J. Geophys. Res. Atmos.* 118, 5380–5552, <https://doi.org/10.1002/jgrd.50171>, 2013.
- Cao, P., Xu, K., Varghese, J. O., and Heath, J. R.: The microscopic structure of adsorbed water on hydrophobic surfaces under ambient conditions, *Nano Lett.* 11, 5581–5586, 2011.
- 340 Dalirian, M., Ylisirniö, A., Buchholz, A., Schlesinger, D., Ström, J., Virtanen, A., and Riipinen, I.: Cloud droplet activation of black carbon particles coated with organic compounds of varying solubility, *Atmos. Chem. Phys.*, 18, 12477–12489, <https://doi.org/10.5194/acp-18-12477-2018>, 2018.
- Defay, R., Prigogine, I., and Bellemans, A.: *Surface Tension and Adsorption*, Longmans, London, 1966.
- Demirdjian, B., Ferry, D., Suzanne, J., Popovicheva, O.B., Persiantseva, N.M., and Shonija, N.K.: Heterogeneities in the
345 microstructure and composition of aircraft engine combustor soot: impact on the water uptake, *J. Atmos. Chem.* 56, 83–103, 2007.
- Dusek, U., Reischl, G. P., and Hitzenberger, R. CCN activation of pure and coated carbon black particles, *Environ. Sci. Tech.* 40, 1223–1240, <https://doi.org/10.1021/es0503478>, 2006.
- Fletcher, N.H.: Size effect in heterogeneous nucleation. *J. Chem. Phys.* 29, 572–576, 1958.
- 350 Franks, F.: Nucleation of ice and its management in ecosystems, *Phil. Trans. R. Soc. Lond. A* 361, 557–574, 2003.
- Fowkes, F. M., and Harkins, W. D.: The state of monolayers adsorbed at the interface solid-aqueous solution, *J. Am. Chem. Soc.* 62, 3377–3386, 1940.
- Frenkel, J.: *Kinetic Theory of Liquids*, Oxford University Press, London, 1946.
- 355 Hagen, D. E., Trueblood, M. B., and White, D. R., Hydration properties of combustion aerosols, *Aerosol Sci. Tech.* 10, 63–69, 1989.



- Halsey, G., Physical Adsorption on non-uniform surfaces, *J. Chem. Phys.* 16, 931–937, 1948.
- Henson, B. F.: An adsorption model of insoluble particle activation: Application to black carbon, *J. Geophys. Res.* 112, D24S16, <https://doi.org/10.1029/2007JD008549>, 2007.
- 360 Hill, T. L.: Statistical mechanics of adsorption. V. Thermodynamics and heat of adsorption, *J. Chem. Phys.* 17, 520–535, 1949a.
- Hill, T. L., Physical adsorption and the free volume model for liquids, *J. Chem. Phys.* 17, 590 1949b.
- Hrubý, J., Vinš, V., Mareš, R., and Kalová, J.: Surface tension of supercooled water: no inflection point down to -25 °C, *J. Phys. Chem. Lett.* 5, 425–428, 2014.
- 365 Ikhenazene, R., Pirim, C., Noble, J. A., Irimiea, C., Carpentier, Y., Ortega I. K., Ouf, F.-X., Focsa, C., and Chazallon, B.: Ice nucleation activities of carbon-bearing materials in deposition mode: From graphite to airplane soot surrogates, *J. Phys. Chem. C* 124, 489–503, 2020.
- Kotzick, R., Panne, U., and Niessner, R.: Changes in condensation properties of ultrafine carbon particles subjected to oxidation by ozone, *J. Aerosol Sci.* 28, 725–735, 1997.
- 370 Kumar, P., Sokolik, I. N., and Nenes, A.: Measurements of cloud condensation nuclei activity and droplet activation kinetics of fresh unprocessed regional dust samples and minerals, *Atmos. Chem. Phys.*, 11, 3527–3541, <https://doi.org/10.5194/acp-11-3527-2011>, 2011.
- Kuznetsov, B. V., Rakhmanova, T. A., Popovicheva, O. B., and Shonija, N. K.: Water adsorption and energetic properties of spark discharge soot: specific features of hydrophobicity, *J. Aerosol Sci.* 34, 1465–1479, 2003.
- 375 Laaksonen, A.: A unifying model for adsorption and nucleation of vapors on solid surfaces, *J. Phys. Chem. A* 119, 3736–3745, 2015.
- Laaksonen, A., and Malila, J.: An adsorption theory of heterogeneous nucleation of water vapour on nanoparticles, *Atmos. Chem. Phys.* 16, 135–143, 2016.
- 380 Laaksonen, A., Malila, J., Nenes, A., Hung, H.-M., and Chen, J.-P.: Surface fractal dimension, water adsorption efficiency, and cloud nucleation activity of insoluble aerosol, *Sci. Rep.* 6, 25504, 2016.



- Lupi, L., Kastelowitz, N., and Molinero, V.: Vapor deposition of water on graphitic surfaces: formation of amorphous ice, bilayer ice, ice I, and liquid water, *J. Chem. Phys.* 141, 18C508, 2014.
- Mahata, P. C., and Alofs, D. J., Insoluble condensation nuclei: the effect of contact angle, surface roughness and adsorption, *J. Atmos. Sci.* 32, 116–122, 1975.
- 385
- Marculli, C.: Pre-activation of aerosol particles by ice preserved in pores, *Atmos. Chem. Phys.*, 17, 1595–1622, <https://doi.org/10.5194/acp-17-1595-2017>, 2017.
- Murphy, D. M., and Koop, T.: Review of the vapour pressures of ice and supercooled water for atmospheric applications, *Q. J. R. Meteorol. Soc.* 131, 1539–1565, 2005.
- 390
- Möller, D.: On the history of the scientific exploration of fog, dew, rain and other atmospheric water, *Erde* 139, 11–44, 2008.
- Oubal, M., Picaud, S., Rayez, M.-T., and Rayez, J.-C.: Water adsorption on oxidized single atomic vacancies present at the surface of small carbonaceous nanoparticles modeling soot, *ChemPhysChem* 11, 4088–4096, 2010.
- Petters, M. D. and Kreidenweis, S. M.: A single parameter representation of hygroscopic growth and cloud condensation nucleus activity, *Atmos. Chem. Phys.*, 7, 1961–1971, <https://doi.org/10.5194/acp-7-1961-2007>, 2007.
- 395
- Popovicheva, O. B., Persiantseva, N. M., Tishkova, V., Shonija, N. K., and Zubareva, N. A.: Quantification of water uptake by soot particles, *Environ. Res. Lett.* 3, 025009, 2008.
- Pruppacher, H. R., and Klett, J. D.: *Microphysics of Clouds and Precipitation*, Kluwer Academic, Dordrecht, ed. 2, 1997.
- Rowley, H. H., and Innes, W. B.: Relationships between the spreading pressure, adsorption, and wetting, *J. Phys. Chem.* 46, 694–705, 1942.
- 400
- Samseth, B. H., Stern, C. W., Andrews, E., Kahn, R. A., Myhre, G., Schulz, M., and Schuster, G. L.: Aerosol absorption: Progress towards global and regional constraints, *Curr. Clim. Change Rep.* 4, 65–83, 2018.
- Seinfeld, J. H., Bretherton, C. S., Carslaw, K. S., Coe, H., DeMott, P. J., Dunlea, E. J., Feingold, G., Ghan, S. J., Guenther, A. B., Kahn, R. A., Kracunas, I. P., Kreidenweis, S. M., Molina, M. J., Nenes, A., Penner, J. E., Prather, K. A., Ramanathan, V., Ramaswamy, V., Rasch, P. J., Ravishankara, A. R., Rosenfeld, D., Stephens, G., and Wood, R.:
- 405



- Improving our fundamental understanding of the role of aerosol-cloud interactions in the climate system, *Proc. Natl. Acad. Sci. U.S.A.* 113, 5781–5790, 2016.
- Sergi, D., Schocci, G., and Ortona, A.: Molecular dynamics simulations of the contact angle between water droplets and graphite surfaces, *Fluid Phase Equilibr.* 332, 173–177, 2012.
- 410 Sorjamaa, R., and Laaksonen, A., The effect of H₂O adsorption on cloud drop activation of insoluble particles: a theoretical framework, *Atmos. Chem. Phys.* 7, 6175–6180, 2007.
- Tadmor, R.: Line energy and the relation between advancing, receding, and Young contact angles, *Langmuir* 20, 7659–7664, 2004.
- Wittbom, C., Eriksson, A. C., Rissler, J., Carlsson, J. E., Roldin, P., Nordin, E. Z., Nilsson, P. T.,
415 Swietlicki, E., Pagels, J. H., and Svenningsson, B.: Cloud droplet activity changes of soot aerosol upon smog chamber ageing, *Atmos. Chem. Phys.*, 14, 9831–9854, <https://doi.org/10.5194/acp-14-9831-2014>, 2014.
- Young, G. J., Chessick, J. J., Healey, F. H., and Zettlemoyer, A. C., Thermodynamics of the adsorption of water on graphon from heats of immersion and adsorption data, *J. Phys. Chem.* 58, 313–315, 1954.
- Zheng, Y., and Zaoui, A.: Wetting and nanodroplet contact angle of the clay 2:1 surface: The case
420 of Na-montmorillonite (001), *Appl. Surf. Sci.* 396, 717–722, 2017.