Referee #1

Dear Referee,

Thank you very much for you thoughtful and constructive comments for our work. We have addressed them all in the revised version of the manuscript as described below:

In this work, Nijing Wang and colleagues presented ship-borne measurements of a number of atmospheric carbonyl compounds around the Arabian Peninsula. During the AQABA campaign in 2017, a comprehensive suite of scientific payload provided the simultaneous measurements of an array of trace gases, including the carbonyl compounds measured using a PTR-MS system. The impacts of oceanic emissions and the oil and gas industry in this region are discussed. The manuscript is aligned with the scope of ACP, and the topic is of interest to the community.

A clear focus is given to acetaldehyde in this manuscript. It has been recognized that in situ measurement techniques of acetaldehyde may be subject to substantial inlet artifacts, especially in remote environments. Such artifacts may be related to tropospherically relevant ozone levels and have been reported for the PTR-MS system on research aircraft platforms (e.g., Northway et al., 2004). Compared to airborne measurements, in situ instruments onboard research vessels usually require much longer sampling lines due to logistic reasons; and the ship may had frequently encountered organic-rich air masses (e.g., polluted air, ship exhaust, and marine air with sea spray aerosols and possibly unsaturated organic compounds). The potential inlet artifacts for the ship-borne trace gas measurements have not been thoroughly discussed in many previous ship-borne studies. In this work, the authors presented fairly detailed results and discussion regarding the potential inlet artifacts, and concluded that the inlet artifacts, although cannot be fully ruled out, are unlikely to be significant in this dataset. I do appreciate the efforts the authors have invested in the potential inlet artifacts. Based on the results presented in this manuscript, I cautiously agree with the authors that the inlet artifacts are probably not a huge concern in this work. More thorough inlet tests would provide more information, which should be considered in future cruise deployments.

We thanks the reviewer for noting and agreeing with our assessment of the potential inlet artifact. Through this work we have become more aware of this issue and will characterize the inlet accordingly in future measurement campaigns.

In addition, a global chemical transport model (EMAC) is used to examine the sources and sinks of these compounds in this region. The modeling component of this work certainly provides valuable insights. A more thorough model evaluation should also be carried out for some of the key precursors of these carbonyl compounds, such as ethane, propane, butanes, etc.

A thorough model evaluation of the major alkanes measured during AQABA has been published in *Nature Communications* earlier this year (Bourtsoukidis et al., 2020)., The aforementioned study also included multiple sensitivity analyses by the model to investigate alkane sources. We therefore refer to the published work in this manuscript to avoid duplication. . We have further addressed the model comparison results in terms of carbonyl precursors in section 3.4.3.

However, the empirical analysis presented in Section 3.2.1 is largely based on fundamentally flawed assumptions, especially for formaldehyde, acetaldehyde, and acetone. I will elaborate this later. This is one major drawback of this study. For this reason, it is my opinion that major revision is needed before publication in ACP. I would recommend the authors remove this section entirely. Or, perhaps some sort of box modeling (e.g., diurnal steady state model constrained to observed hydrocarbons and other measurements available) may be useful.

On reflection, we agree with the reviewer that this section did not contribute significantly to the analysis and we decided to remove this section entirely.

Detailed response:

Page 1, line 14: "In this study we examine carbonyl compounds (CxHyO), …" Clearly the formula CxHyO can be an alcohol, an ether, or an epoxide. I do not think a general formula is useful here.

We agree. CxHyO is removed in the text.

Page 2, line 47: "…important source of free radicals (HOx)" should be hydroxyl radicals here. Also x should be in subscript. Unless the authors have other radicals in mind, in which case it should be more specific.

We revised to "hydroxyl and hydroperoxyl radicals free radicals (HO_x)" in the text.

Page 2, line 48: NOx: please define all abbreviations the first time it appears in the main text. Also x should be in subscript.

We revised to "NO_x (NO and NO_2)" in the text.

Page 2, line 56-57: for the sake of completeness, should cite a few more previous studies here (e.g., Yang et al., 2014; Wang et al., 2019), in which tens of ppt of acetaldehyde were reported in the remote troposphere.

More references were added in the text:

Remote area:

Wisthaler, A.: Organic trace gas measurements by PTR-MS during INDOEX 1999, Journal of Geophysical Research, 107, 10.1029/2001jd000576, 2002.

Yang, M., Beale, R., Liss, P., Johnson, M., Blomquist, B., and Nightingale, P.: Air–sea fluxes of oxygenated volatile organic compounds across the Atlantic Ocean, Atmospheric Chemistry and Physics, 14, 7499-7517, 10.5194/acp-14-7499-2014, 2014.

Wang, S., Hornbrook, R. S., Hills, A., Emmons, L. K., Tilmes, S., Lamarque, J. F., Jimenez, J. L., Campuzano ‐ Jost, P., Nault, B. A., Crounse, J. D., Wennberg, P. O., Kim, M., Allen, H., Ryerson, T. B., Thompson, C. R., Peischl, J., Moore, F., Nance, D., Hall, B., Elkins, J., Tanner, D., Huey, L. G., Hall, S. R., Ullmann, K., Orlando, J. J., Tyndall, G. S., Flocke, F. M., Ray, E., Hanisco, T. F., Wolfe, G. M., St. Clair, J., Commane, R., Daube, B., Barletta, B., Blake, D. R., Weinzierl, B., Dollner, M., Conley, A., Vitt, F., Wofsy, S. C., Riemer, D. D., and Apel, E. C.: Atmospheric Acetaldehyde: Importance of Air ‐ Sea Exchange and a Missing Source in the Remote Troposphere, Geophysical Research Letters, 10.1029/2019gl082034, 2019.

Polluted area:

Koss, A. R., de Gouw, J., Warneke, C., Gilman, J. B., Lerner, B. M., Graus, M., Yuan, B., Edwards, P., Brown, S. S., Wild, R., Roberts, J. M., Bates, T. S., and Quinn, P. K.: Photochemical aging of volatile organic compounds associated with oil and natural gas extraction in the Uintah Basin, UT, during a wintertime ozone formation event, Atmospheric Chemistry and Physics, 15, 5727-5741, 10.5194/acp-15-5727-2015, 2015.

Page 2. Line 59: "… in those regions" this can be confusing. Please revise.

We revised to “Generally, secondary photochemical formation from various precursors is the main source for those carbonyl compounds.” in the text to avoid the confusion.

Page 4, line 100: VOC container might be misleading here. I’m guessing this is one of the lab containers loaded aboard, but it could be confused with some sort of container for volatile organic compounds.

We revised “VOC container” to “on-board lab container” in the text.

Page 5, line 144: please provide a bit more info about this spectral radiometer, such as manufacturer, model, etc.

The manufacturer and related reference were added in the text as “The actinic flux was measured by a spectral radiometer (Metcon GmbH; Meusel et al., 2016).”

Page 5, line 145: since NMHCs provide vital constraints on the budget analysis in this work, please provide a bit more information here. For example, is this an online or canister-based GC-FID?

More information about NMHCs measurement was added in the text as “Non methane hydrocarbons (NMHC) mixing ratios were measured by a gas chromatograph with flame ionization detector (GC−FID) online with the time resolution of 50 minutes. It measured hydrocarbons (C2 - C8) and aromatics (C6 - C8) with the average LOD < 10 ppt for most of compounds. For a detailed instrumental description see Bourtsoukidis et al. (2019).”

Page 5, line 156: I understand that EMAC is a well established model and many key components are archived elsewhere. Given the importance of oceanic emissions in this work, I am sure the readers would appreciate a bit more details, especially how the oceanic emissions are setup/configured for the compounds discussed in this work, such as acetaldehyde, acetone, etc.

More information on EMAC and in particular the air-sea transfer was added in the text as follows “… a comprehensive chemistry mechanism MOM (Mainz Organic Mechanism) was deployed (Sander et al., 2019). The model considers direct emissions (such as anthropogenic, biogenic, biomass burning etc.), atmospheric transport and mixing, photochemical production of carbonyls (by OH, O_3 and NO_3), as well as physical and chemical removal processes. The exchange of organic compounds between ocean and atmosphere was considered in EMAC via the AIRSEA submodel, described in detail in Pozzer et al. (2006). The transfer velocity is calculated online and the concentration in the water is prescribed by the user. For acetone, a constant water concentration of 15 nmol/L is used, following the suggestion of Fischer et al. (2012), while direct oceanic emission of acetaldehyde was initially not included in the model. The model configuration in the study is the same as the model applied in Bourtsoukidis et al. (2020) in the resolution of T106L31 (i.e. ~ $1.1^{o} \times 1.1^{o}$ horizontal resolution and , 31 vertical hybrid pressure levels up to 10 hPa) and the time resolution of 10 minutes. The measurement data of PTR-ToF-MS were averaged to 10-minute resolution to match the model data resolution for further comparison.”

Page 8, line 210-223: are these carbonyl compounds (i) directly emitted from the oil & gas facilities; (ii) produced from precursors emitted from these facilities; or (iii) associated with other anthropogenic activities in this area?

We discussed these sources (first or secondary formation of carbonyls from anthropogenic activities) in section 3.1.2 (case studies of polluted regions) and also section 3.4.3 (anthropogenic primary sources for underestimated acetaldehyde). In order to make this clear, we now added a sentence in the end of this paragraph as follow: The sources of the major carbonyls in the Arabian Gulf will be discussed in details in section 3.1.2 and 3.4.3.

Page 10, Table 2: I am not entirely sure about the purpose of this table. I understand that the overarching goal here is to put the measurements obtained in this study into the context of previous studies. However, this table itself certainly does not cover the comprehensive characteristics of any of the category. For example, some of the studies cited in the marine category contains clear influence from anthropogenic emissions/ biomass burning; none of the cited studies in this category reflects the pristine marine environments, such as tropics, Southern Ocean, etc. For this reason, this table may create wrong impression to the audience. I understand that this is not a review paper, but a more thorough review of the literature is always appreciated. Therefore, I would suggest that the authors compile a slightly more exhaustive list, and be cautious when quoting/discussing the numbers in the literature. I myself find some of the categories are less relevant for this work (e.g., urban, rural, forest), and the marine category can certainly use more details (e.g., coastal vs open ocean, northern vs southern hemisphere, high latitude oceans, natural vs influenced by anthropogenic/biomass burning emissions). I would also recommend the authors add relevant info (e.g., lat/lon) to this table.

We now updated the table by including 5 more studies done in open sea and coastal areas. We replaced the study of oil and gas influenced regions by other four studies. Studies done in urban, rural and forest regions were removed from the table. We included the coordinates (latitude and longitude) and measurement sampling height for those studies as requested. Accordingly, we updated the text related to the comparison of previous studies in the revised manuscript.

Page 9, line 236-237: why do the authors compare the ship-borne measurements obtained in this work to the measurements from a rural site in Cyprus? The numbers may be similar for vastly different reasons. Or do the authors imply a connection with Cyprus? If yes, please clarify & elaborate.

As suggested by the reviewer, the table has been updated and the measurements results from this Cyprus study as well as the related text were removed and rewritten in the revised manuscript as: "… The Mediterranean Sea had somewhat higher levels of aliphatic carbonyls than the clean regions (the Arabian Sea and the Gulf of Aden) but with acetone (above 2ppb) dominating the distribution. Much higher acetone level than acetaldehyde level was also observed for some costal site measurement which was impacted by continental air (White et al., 2008; Schlundt et al., 2017, see Table 2). …"

Page 9, line 247-248: note that acetone may deposit into the ocean, especially in the continental outflow from polluted regions (e.g., Marandino et al., 2005; Schlundt et al., 2017; Yang et al., 2014).

We are aware that under conditions of polluted continental outflow over the ocean uptake of acetone from the air to the sea will occur. Also dilution and mixing with free tropospheric air during transport can modulate acetone mixing ratios. This process is accounted for in the EMAC model used in the subsequent analysis. We now added a sentence in the revised manuscript to make it clear as follow: "… could be also transported from the Mediterranean Sea (where acetone was high). Four-day back trajectories indicate the air reaching Suez region was mostly originated from Europe continent passing over the Mediterranean Sea (Bourtsoukidis et al., 2019). Meanwhile, ocean uptake of acetone from the air due to polluted continental outflow (Marandino et al., 2005) as well as dilution and mixing with free tropospheric air during transport can modulate acetone mixing ratios."

Page 9, line 253: Again, the authors compare the results in Red Sea to that from the other side of the planet (Thompson Farm, a rural site in the US). What is the point of the seemingly random comparison?

With these comparisons we were trying to set the AQABA measurements in a broader context. Since this appears to have been a distraction, we have followed the reviewer's suggestion, and the table has been updated, so that the measurements from the Thompson Farm study as well as the related text were removed in the revised manuscript as follows:"…The mixing ratios of acetaldehyde and acetone over the northern part of the Red Sea were similar to those levels measured in western Pacific coastal regions (South China Sea, Table 2)…."

Page 11, line 269: are these numbers the sum of all measured unsaturated and aromatic carbonyls?

The numbers represent the range of the mean mixing ratios of unsaturated carbonyls in each region. To make this clear we added the sentence "…followed by Suez (11 – 68 ppt). The numbers represent the range of the mean mixing ratios of unsaturated carbonyls in each region"

Page 11, line 293-295: these very general descriptions belong to the introduction section.

These lines were removed in the revised manuscript.

Page 11, line 298-299: The first two assumptions do not hold, for the following reasons: (1) Photolysis can be a significant sink for certain carbonyl compounds, such as acetone (Fischer et al., 2012); for formaldehyde it's even the dominant sink (Anderson et al., 2017). (2) Primary emissions contribute substantially to many carbonyl compounds in the atmosphere, such as acetone (e.g., Fischer et al., 2012), acetaldehyde (e.g., Millet et al., 2010). Biomass burning contribution to many of these carbonyl compounds is also substantial, and the authors even discussed the impacts of biomass burning in Section 3.2.2. Therefore, primary emissions certainly cannot and should not be ignored. (3) Dilution (mixing with background air) can lead to rapid decrease of tracer concentrations in freshly emitted plumes. Uptake by the terrestrial vegetation and the ocean may also affect the carbonyl budget on the regional and global scale. The impact of mixing may be cancelled out by scaling the mixing ratios of hydrocarbons and carbonyls to a long-lived compound, such as CO or ethylene, but I cannot think of a simple way to overcome other impacts as I listed above. In addition to the flawed assumptions, the carbonyl yields are not constant and depend on other variables (e.g., NOx levels, oxidation pathways). Table S4 did not clarify how the yields were derived.

Thanks for pointing out the inappropriate application of the empirical method we applied in our study. The intention was to show what the photochemistry alone (without these influences) would generate. We agree that because of the complexity of multiple alternate sources this section does not have much scientific contribution to the whole manuscript. As we compared the measurement with a comprehensive global model in the manuscript (which includes such sources), we removed the part regarding the empirical calculations and corrected relevant text in the revised manuscript. For the yield in Table S4, we assumed the alkylperoxy radicals (RO_2) would mainly react with NO (high NO condition) and then summed up the yield of all the reaction pathways of hydrocarbons that could produce carbonyls. Now as we decided to remove the empirical calculation section, the Table S4 was also removed from the supporting information.

Page 12, line 314-319: This exercise is only valid when the ratio in the source region is well understood, which is clearly not the case here. Figure S3 cannot rule out the possibility that there may be air masses with different initial toluene/benzene ratios mixed together. It certainly looks like there are multiple clusters in Figure S3. This comes back to my previous question regarding the oil and gas industry in this region: do these measurements reflect (i) direct emissions from the oil & gas facilities; or (ii) co-located (anthropogenic) emissions in this area? This type of photochemical clock does not work without a thorough understanding of the source characteristics. To sum up, the OH exposure calculation may be useful, provided that the source contributions are well

understood. The empirical analysis based on Equation 1 and 2, however, does not really provide scientifically valuable information, definitely not for acetone, acetaldehyde, and formaldehyde.

We agree that it is not appropriate to use the hydrocarbon ratios to calculate the OH exposure for the remote areas. However, we kept the estimations of OH exposure using hydrocarbon ratios for the polluted regions (Arabian Gulf and Suez) where the major primary emissions were identified based on NMHCs data reported by Bourtsoukidis et al., (2019) and Bourtsoukidis et al. (2020). We now mention this specifically in the section 3.1.2 in the revised manuscript and also compared the toluene to benzene emission ratios to previous studies having similar sources. The revised text is as follows:

…For further insight, we focused on a time series of selected trace-gases and their inter-correlations to better identify the sources of the major aliphatic carbonyls. Meanwhile, we calculated the OH exposure ($[OH]\Delta t$) based on hydrocarbon ratios (Roberts et al., 1984;de Gouw et al., 2005;Yuan et al., 2012) for the polluted regions Arabian Gulf and Suez where primary emissions have been identified (Bourtsoukidis et al., 2019; Bourtsoukidis et al. 2020), to better understand the photochemical aging of the major carbonyls using the following equation:

$$[OH]\Delta t = \frac{1}{k_X - k_Y} \cdot \left(In \frac{[X]}{[Y]}\bigg|_{t=0} - In \frac{[X]}{[Y]} \right) , \qquad \text{Eq. (1)}$$

where X and Y refer to two hydrocarbon compounds with different rates of reaction with the OH radical (k). For this study, we chose toluene ($k_{OH+toluene}$: 5.63E-12 cm^3 $molecule^{-1}s^{-1}$) and benzene ($k_{OH+benzene}$: 1.22E-12 cm^3 $molecule^{-1}s^{-1}$) (Atkinson and Arey, 2003), because both compounds were measured by PTR-ToF-MS at high frequency and these values showed a good agreement with values measured by GC-FID (Figure S2). The approach detailed by Yuan et al. (2012) was applied to determine the initial emission ratio $\frac{[X]}{[Y]}\Big|_{t=0}$ in those two regions by only including nighttime data of benzene and toluene. We obtained initial emission ratios (toluene to benzene ratio) of 1.38 for the Arabian Gulf and 2.12 for the Suez region. Koss et al. (2017) summarized the toluene to benzene ratios observed in various locations and showed that urban and vehicle sources tend to have higher toluene to benzene ratio (mean ~ 2.5) than the ratios of oil & gas sources (mean ~ 1.2). Therefore, the toluene to benzene ratios obtained for those two regions agreed well with other studies done with similar emissions sources. The corresponding correlation plots of toluene and benzene for those two regions can be found in Figure S3.

And we removed this description to the case studies because we included the OH exposure to help discuss the air mass age.

Page 14, Figure 4: Units for j_NO2 and OH exposure are missing. Same for Figure 5.

The units were added in the figures of the revised manuscript.

Page 14, line 366: "As mentioned before, photochemical oxidation contributed a large fraction to acetone and the larger aliphatic carbonyls over the Arabian Gulf and Suez areas" I think this statement needs stronger support. The terrestrial biosphere may emit a large amount of acetone. Moreover, the acetone yield from terpene oxidation is quite high, and the acetone production from terpenes may not be a huge source globally, but could well be in certain regions. Similarly, some of these compounds may be from anthropogenic sources or biomass burning.

As the empirical calculation was removed, this conclusion previously derived from the results were removed as well. However, we would like to address the comment given by the reviewer as follows. The high mixing ratios of acetone precursors (propane and several higher alkanes) combined with strong photochemical processing, provide a strong regional secondary source. Regional biogenic sources of acetone are likely to be small given that much of the Arabian Peninsula is desert and little isoprene (a common biogenic emission) and monoterpenes was detected (most of the time below detection limit and the average was around tens of ppt). In the Arabian Gulf no significant increase of acetonitrile was observed so that a biomass burning contribution to acetone can be ruled out. In the region of Suez, acetonitrile levels were somewhat elevated. The lowest of acetone to acetonitrile ratio was only during the biomass burning plum spikes in Suez, ranging from 7-10, which was similar to the ratios of acetone to acetonitrile reported by Holzinger et al. (2005) in aged biomass burning plume over the Eastern Mediterranean. During the rest of the time, the acetone to acetonitrile ratio was above 10 and up to 30. It suggested that the biomass burning source is still likely minor compared to the secondary source.

Holzinger, R., Williams, J., Salisbury, G., Klüpfel, T., de Reus, M., Traub, M., Crutzen, P. J., and Lelieveld, J.: Oxygenated compounds in aged biomass burning plumes over the Eastern Mediterranean: evidence for strong secondary production of methanol and acetone, Atmos. Chem. Phys., 5, 39-46, 10.5194/acp-5-39-2005, 2005.

Page 14, line 374-376: The authors concluded that the carbonyls are co-produced via photochemical oxidation because they show strong correlation with ozone. I am not totally convinced. It is plausible that these carbonyls may be co-emitted with ozone precursors.

We have added support for this statement by referencing the work of Tadic et al., (2020), who studied the production of O3 in the region in detail. In order to retain the reviewer's point we also note that primary emissions may also occur, as follows "Tadic et al. (2020) reported the net ozone production rate over the Arabian Gulf (32 ppb d^{-1}) was the greatest over the Arabian Peninsula. They show that strong ozone forming photochemistry occurred in this region, which would lead to abundant secondary photo-chemically produced products (including carbonyls). However, it should be noted the good correlation between ozone and carbonyls could in part be due to carbonyls co-emitted with ozone precursors (hydrocarbons) as primary emissions."

Tadic, I., Crowley, J. N., Dienhart, D., Eger, P., Harder, H., Hottmann, B., Martinez, M., Parchatka, U., Paris, J.-D., Pozzer, A., Rohloff, R., Schuladen, J., Shenolikar, J., Tauer, S., Lelieveld, J., and Fischer, H.: Net ozone production and its relationship to nitrogen oxides and volatile organic compounds in the marine boundary layer around the Arabian Peninsula, Atmospheric Chemistry and Physics, 20, 6769-6787, 10.5194/acp-20-6769-2020, 2020.

Page 14, line 385: I noticed that the spatial distribution oil fields and refineries is from the year of 2007, while this cruise campaign was conducted in 2017. Is it possible to obtain more recent information on this?

Unfortunately, we do not have access to any newer dataset of the oil and refineries distributions. Previously published work regarding the AQABA (Bourtsoukidis et al., 2019; Pfannerstill et al., 2019) all reported the distribution from 2007 data base.

Page 15, line 405: "The biomass burning emissions were probably transported by on the prevailing northerly wind above Northeast Egypt…" is this supported by backward trajectory analysis or meteorological conditions?

The prevailing wind direction in Suez was shown in Figure S1. We added "…by on the prevailing northerly wind (Figure S1) above Northeast Egypt…" in the revised manuscript.

Page 16, Figure 6: this is interesting. how are these not-assigned ions calibrated/quantified?

Where no calibration gas is available but a mass is detected then the mixing ratios of those ions were calculated based on an established theoretic calculation method (Lindinger et al. 1998) using a fixed proton transfer reaction rate constant (k_{PTR}) of 2.0 × 10^{-9} cm^3 s^{-1}. Fortunately the protonation rate constant for all compounds is very similar, so that this approach yields reasonable results when direct in-field calibration is not possible.

Lindinger, W., Hansel, A., and Jordan, A.: On-line monitoring of volatile organic compounds at pptv levels by means of proton-transfer-reaction mass spectrometry (PTR-MS) medical applications, food control and environmental research, International Journal of Mass Spectrometry and Ion Processes, 173, 191-241, 1998.

Page 17, line 449-452: any GC-FID measurements of isoprene during this period?

Yes. The GC-FID also measured isoprene during AQABA campaign. However, as mentioned already in the text (line 417-418): According to the GC-FID measurement, isoprene was below the detection limit for most of the time during the AQABA cruise with the highest values observed in Suez (10 - 350 ppt).

Page 17, line 460-462: "The model considers direct emissions (such as anthropogenic, biogenic, biomass burning etc.), atmospheric transport and mixing, photochemical production of carbonyls (by OH, O3 and NO3), and physical and chemical removal processes." This belongs to the method section where the model is introduced.

We now moved this part to the method section 2.5.

Page 17, line 465-466: "acetone was overestimated by a factor within 1.5 over the Arabian Sea, Gulf of Aden and Gulf of Oman, and underestimated by a factor within 2.5 over the other regions." Note that previous studies reported downward oceanic fluxes of acetone (ocean being a net sink) in the continental outflow from polluted regions (e.g., Marandino et al., 2005; Schlundt et al., 2017; Yang et al., 2014). How does the model treat the air-sea exchange of acetone? This key information is missing.

The sub-model AIRSEA (Pozzer et al. 2006) implemented in EMAC calculates the exchange of acetone between the ocean and the atmosphere. The transfer velocity is calculated online. In general, ocean is a net sink for acetone, but regionally could be an emitter of acetone. This strongly depends on its concentration, however, away from sources the surface seawater is close to equilibrium with boundary layer air (Williams et al. 2004). For acetone in the model, a constant water concentration of 15 nmol/L is used, following the suggestion of Fischer et al. (2012). We added this information in the method part (section 2.5).

Fischer, E. V., Jacob, D. J., Millet, D. B., Yantosca, R. M., and Mao, J.: The role of the ocean in the global atmospheric budget of acetone, Geophysical Research Letters, 39, n/a-n/a, 10.1029/2011gl050086, 2012.

Pozzer, A., Jöckel, P., Sander, R., Williams, J., Ganzeveld, L., and Lelieveld, J.: Technical Note: The MESSy-submodel AIRSEA calculating the air-sea exchange of chemical species, Atmos. Chem. Phys., 6, 5435-5444, 10.5194/acp-6-5435-2006, 2006.

Williams, J., Holzinger, R., Gros, V., Xu, X., Atlas, E., and Wallace, D. W. R.: Measurements of organic species in air and seawater from the tropical Atlantic, Geophysical Research Letters, 31, 10.1029/2004gl020012, 2004.

Page 19, line 504-505: "This indicates that the source of acetaldehyde was probably not from direct biogenic production…" this sentence is confusing. I think "direct oceanic acetaldehyde emissions are probably insufficient to explain…" might be more accurate?

Here we wanted to emphasize that acetaldehyde is more related to a non-biogenic emission mechanism based on its poor correlation with DMS. We now correct the sentence to make it clearer: "This indicates that the direct biogenic acetaldehyde emissions from the ocean are probably insufficient to explain the measured acetaldehyde."

Page 19, line 513-514: "To our knowledge, there is no clear experimental evidence showing the ocean to be a sink for acetaldehyde." This is probably true, but the language is vague (I think "direct" is a better choice than "clear" here), and several important studies are missing here. Schlundt et al. (2017) reported net downward fluxes of acetaldehyde in the polluted marine boundary layer (ocean is a net sink), which is inferred from measurements in the atmosphere and the surface ocean (so not "direct" evidence). Yang et al. (2014) reported oceanic fluxes of acetaldehyde using eddy covariance method (this is direct flux measurements). Indeed the fluxes were mostly upward for acetaldehyde (ocean is a net source), there appears to be a few points indicative of net downward fluxes, although are probably close to the limit detection of that system.

Thanks for pointing out that several important studies were omitted when considering this point. We have added now them in the revised manuscript as follows: "… reporting the upper limit of the net ocean emission of acetaldehyde to be 34 Tg a^{-1}. Yang et al. (2014) quantified the air-sea fluxes of several OVOCs over Atlantic Ocean by eddy covariance measurements, showing ocean is a net source for acetaldehyde. Although Schlundt et al. (2017) reported uptake of acetaldehyde by the ocean from measurement-inferred fluxes in western Pacific coastal regions, to our knowledge, there is no direct experimental evidence showing the ocean to be a sink for acetaldehyde. "

Page 21, line 549-554: can the authors please provide new plots like Figure 7 but for other major acetaldehyde/acetone precursors, such as ethane, propane, butanes, and perhaps ethanol as well? This would certainly make the case stronger, and may provide key insight into the acetaldehyde budget in this region. Even "remote" regions in this work are not really that far from the source regions. Therefore it remains unclear to what degree the underestimation/overestimation of carbonyls is due to their precursors. Emission inventories often have a hard time capturing the anthropogenic emissions, especially from the oil and gas industry.

Exactly these precursor plots are already provided in the paper of Bourtsoukidis et al. (2019). Furthermore, Bourtsoukidis et al. (2020) already published a measurement-model comparison for the key precursors (ethane, propane and butane). After implementing the new deep water source, the model was able to mostly reproduce the measurements of hydrocarbons over most areas except for the large underestimation over Suez and the Arabian Gulf, which was mentioned in the text. Therefore, we think it likely that other precursors which were not included in the model may contribute to the underestimation of acetaldehyde. Ethanol was not measured during the campaign. We do now include model comparison of alkenes (ethene and propene) to further support our argument of nighttime ozonolysis as a potential source in the Arabian Gulf as suggested by the other referee. In order to more clearly demonstrate the anthropogenic contribution to the model bias of acetaldehyde, we revised manuscript as follows:

Over the Arabian Gulf and Suez, the intensive photochemical production of carbonyls is apparent. Bourtsoukidis et al. (2020) compared measured hydrocarbons (ethane, propane, and butane) with the results from model simulations (the same model used in this study with the newly discovered deep water source implemented). The model was able to reproduce the measurement over most regions except for some significant model underestimations in Suez and Arabian Gulf, in which local and small-scale emissions were difficult for the model to capture. Therefore, an underestimation of the precursor hydrocarbons, as well as those large alkanes, alkenes and cyclic hydrocarbons which were not measured (> C8) or included in the model (> C5), could be a reason for the model underestimation of acetaldehyde especially in polluted regions. In addition, as mentioned in the previous case studies, high ozone mixing ratios were observed over the Arabian Gulf especially during the nighttime. Ethene and propene were found to be significantly underestimated during the nighttime high ozone period by a factor over 10 (Figure S9), which indicates that the nighttime ozonolysis of alkenes could be another important source for acetaldehyde, formaldehyde and other carbonyls (Atkinson et al., 1995;Altshuller, 1993) in the Arabian Gulf.

Page 21, line 570: what biomass burning emission inventory is used in this work?

For the biomass burning emissions, we used the global fire assimilation system emissions data from Kaiser et al. (2012). A sentence has been added in the method section 2.4 to address: The global fire assimilation system was used for biomass burning emissions (Kaiser et.al., 2012)."

Kaiser, J. W., Heil, A., Andreae, M. O., Benedetti, A., Chubarova, N., Jones, L., Morcrette, J.-J., Razinger, M., Schultz, M. G., Suttie, M., and van der Werf, G. R. (2012). Biomass burning emissions estimated with a global fire assimilation system based on observed fire radiative power. BG, 9:527-554

Referee #2

Dear Referee,

Thank you very much for you thoughtful and constructive comments for our work. We addressed them in the revised version of the manuscript as described below:

Overview:

The manuscript presented ship-borne measurements of carbonyl compounds (carbonyls) around the Arabian Peninsula during the AQABA campaign (end of June to end of August 2017) and interpreted the measurement with a chemical transport model (EMAC). Analyses were focused on 3 classified carbonyl groups—aliphatic ($C_nH_{2n}O$), aromatic ($C_nH_{2n-8}O$), and unsaturated ($C_nH_{2n-2}O$)—in terms of their regional distribution, source characterization and partitioning, model-measurement comparison in 8 geographical regions, and finally an in-depth discussion on the missing sources of acetaldehyde (C2 aliphatic carbonyl).

The manuscript is well organized with clear methods, thorough analyses, and coherent arguments. The topic is of great importance given that carbonyls can provide key clues on reactive carbon source and chemistry as an oxygenated and stable group of VOCs, and acetaldehyde has drawn increased attention recently with growing evidence of its large missing sources. This manuscript therefore provided valuable information in advancing the knowledge of atmospheric reactive carbon in the Middle East. I support the publication of the manuscript once the following issues being addressed.

Thank you for noting the key points and significance of the study.

1) **The title is interesting but over-weights the topic towards acetaldehyde and meanwhile lacks a reflection of the large portion of work done on the overview analysis of different carbonyls in this region as presented in the manuscript.**
   On reflection, we agree. The title is now changed to “Measurements of carbonyl compounds around the Arabian Peninsula: overview and model comparison”.

2) **L22. “A smaller but still significant discrepancy …”, are you trying to say that the model-measurement discrepancy of other carbonyls is smaller comparing to that of acetaldehyde? This sentence is not clear and needs to be reorganized.**
   We revised this sentence in the revised manuscript as “We compared the measurements of acetaldehyde, acetone and methyl ethyl ketone to global chemistry-transport model (EMAC) results. A significant discrepancy was found for acetaldehyde, with the model underestimating the measured acetaldehyde mixing ratio by up to an order of magnitude.”

3) **L130. How did you deal with below LOD data?**
   The following statement was added to make this clear: “ … and 9 ± 6 ppt for methyl ethyl ketone (MEK) (Table S1). The data below LOD were excluded from the data set instead of giving zero.”

4) **L134-140. There are other measurement techniques that can differentiate ketone and aldehyde at a shared mass, better to give an average or a range of in-situ reported ketone:aldehyde ratio (e.g., acetone:propanal ratio) to prove the dominance of ketone.**
   We added this information now in the revised manuscript as “… ketones tend to have longer atmospheric lifetimes and higher photochemical yields than aldehydes as mentioned in the introduction. The ratio of measured propanal

to acetone was 0.07 in the western Pacific costal region (Schlundt et al., 2017), 0.06 in an urban Los Angeles (Borbon et al., 2013) and 0.17 - 0.22 in oil &gas production regions (summarized by Koss et al., 2017)…."

5) **Figure2. The description of air mass characteristics is not clear. For example, what's the difference between "source nearby" vs "source nearby, aged", "remote, aged" vs "remote, clean"?**

We agree that the original description may confuse readers. Now in the revised manuscript we have simplified the labeling of the air mass characteristics by dividing them into source nearby (Suez, Arabian Gulf, Gulf of Oman and Red Sea North) and remote (Mediterranean Sea, Red Sea South, Gulf of Aden and Arabian Sea), which is consistent with what Bourtsoukidis et al. (2019) reported based on the NMHCs variability-lifetime results (b factor).

6) **L246-247. "Interestingly, the mean acetaldehyde mixing ratio (0.62 ± 0.59 ppb) over Suez was twice the level found over the Mediterranean Sea, whilst the acetone level was only slightly higher", why is this interesting? What are you trying to imply?**

The main message we would like to give is that the air in Suez region was influenced by local formation in addition to the transportation from Mediterranean Sea, since the mixing ratios do not simply scale, we agree that the original text was not clear. We have now rewritten the text as follows: "Another region where abundant aliphatic carbonyls were observed was Suez region. The air in this region was mainly influenced by nearby cities and marine transportation (ship emissions within the Suez Chanel) (Bourtsoukidis et al., 2019;Pfannerstill et al., 2019). Therefore abundant precursors were available in Suez region, producing more carbonyls regionally especially for shorter-lived compounds (formaldehyde and acetaldehyde). Besides the local-scale emissions and photochemical production contribution to the carbonyls over Suez, the longer lived carbonyls (e.g. acetone) could be also transported from the Mediterranean Sea (where acetone was high). Four-day back trajectories indicate the air reaching Suez region was mostly originated from Europe continent passing over the Mediterranean Sea (Bourtsoukidis et al., 2019)….".

7) **L247-248. "Besides the local-scale emission and photochemical production contribution to the acetone over Suez, the longer lived acetone could be also transported from the Mediterranean Sea (where acetone was high)". Acetone is quite similar in Mediterranean and Suez. "where acetone was high" seems to indicate that acetone source in Suez is dominated by transport rather than local production. We do not know the contribution of transport vs. local source to acetone in Suez region, so need to provide evidence or reorganize the sentence.**

Yes. It is correct that we do not know the contribution fraction of transport and local source to acetone. We reorganized the sentences as specified in the last point (6).

8) **L249-250. "Although the mean mixing ratios of aliphatic carbonyls over Suez were lower than those over the Arabian Gulf, larger variations were observed." I'm not seeing the higher variation over Suez than Arabian Gulf.**

We reorganized the sentence to avoid possible confusion and misunderstanding: "Although the mean mixing ratios of aliphatic carbonyls over Suez were much lower than those over the Arabian Gulf, the variations were still more significant than other regions (not including the Arabian Gulf, see Table 1)."

9) **L254-258. The arguments here are based on back trajectory. How about the proposed deep sea local source of ethane and propane in Bourtsoukidis et al. 2020? This was found not sufficient to explain the model bias**

of acetaldehyde later in the paper, but it worth being brought up here as one potential explanation of high carbonyls in the Red Sea North and also briefly mention what you found later in the model …

We agree that by adding the deep water source information as well as the model comparison result would make the argument more complete. It is now written in the revised manuscript as "…while the sourthern part was more influenced by air from the northern part of the Red Sea mixed with the air masses from desertic areas of central Africa. Therefore, less primary precursors as well as carbonyls were transported to the sourhtern part of the Red Sea compared to the northern part. Moreover, the unexpected sources of hydrocarbons (ethane and propane) from Northern Red Sea deep water reported by Bourtsoukidis et al. (2020) would lead to higher carbonyl levels in the Northen part compared with the Sourthern part due to the additional precursors in the Red Sea North. However, acetaldehdye was still found to be significantly underestimated compared to the model results, even taking the deep-water source into consideration (section 3.3). This indicates that extra sources of acetaldehyde may exist, which will be disscussed in detail in section 3.4."

10) Table 2. Are you summarizing only surface or tower measurement? Need to clarify.

The content of Table 2 was also questioned by reviewer 1. We have now updated the table by including more relevant studies, done in open sea and coastal areas because AQABA measured the marine atmosphere for most of the time. We included the coordinates (latitude and longitude) and measurement sampling height (above sea level) of these studies.

Too few examples of Oil & Gas region and Forest, add more.

We now added more studies done related to Oil&Gas regions. However, although lots of studies reported hydrocarbons (NMHCs) from oil&gas region over the world, we could only find OVOCs results of oil&gas related studies done in the US. In terms of the forest category, we excluded together with other categories (urban and rural) because they were less related to AQABA ship campaign atmosphere.

11) L291, the title of Section 3.2 "Chemistry of aliphatic carbonyls", does not reflect the source characterization discussed in the two case studies in Section 3.2.2.

The section 3.1 and 3.2 were reorganized as we removed one section which will be addressed in the next point. The section titles were changed to
3.1 Aliphatic carbonyls ($C_nH_{2n}O$)
3.1.1 Overview
3.1.2 Case studies of polluted regions: the Arabian Gulf and Suez
3.2 Unsaturated and aromatic carbonyls ($C_nH_{2n-2}O$), ($C_nH_{2n-8}O$)
3.2.1 Overview
3.2.2 Potential precursors and sources of unsaturated carbonyls

12) L303, Eq. (2). Varying formats of this equation have been widely applied but I have one question regarding its application in secondary OVOCs from aged air. Eq. (2) assumes that both precursor hydrocarbon and aliphatic carbonyls experienced the same degree of OH exposure (or Δt). This assumption is reasonable for fresh air mass, but might not be valid for aged air, in which carbonyls kept being produced along the way (depending on its precursor's lifetime) and thus experienced different degrees of OH exposure. So, for long-transported air masses, this could lead to an overestimated OH exposure, an underestimated carbonyl secondary production, and thus an underestimated fractional contribution from oxidation to the measured carbonyl mixing ratio.

Would this uncertainty affect your conclusion (P12, L334-336; P22, L615-618) that hydrocarbon oxidation plays a more important role in polluted regions than other less-polluted regions, given the potential low bias in the estimated contribution for less-polluted/more-remote regions?

Thanks for pointing out the inappropriate application of the Eq. (2) we applied in our study to estimate the carbonyls produced by precursors especially in remote areas where the primary emissions were unknown and the air was aged. The empirical method was also questioned by reviewer 1, because photolysis, primary emissions, dilution effect from the background air and uptake by the land of ocean could influence the small carbonyls to a large extent, factors which were not considered at all in the empirical calculation. We agree that this section does not therefore have much scientific contribution to the whole manuscript. As we compared the measurement with a comprehensive global model in the manuscript, we removed the part regarding the empirical calculations and corrected relevant text in the revised manuscript.

However, we kept the content of estimating the OH exposure using hydrocarbon ratios for the polluted regions where primary emissions were strong and measured (Arabian Gulf and Suez). And we moved this description to the case studies because we included the OH exposure there to help discuss the air mass age.

13) L308-312. Table S4. How is the yield being derived from MCM?

We assumed the alkylperoxy radicals (RO_2) would mainly react with NO (high NO condition) and then sum up the yield of all the reaction pathways of hydrocarbons that could produce carbonyls. Now as we decided to remove the empirical calculation section, the Table S4 was also removed from the supporting information.

14) L318. Did you use 0-5AM data as in Yuan et al. 2012? Clarify that.
L319. How do your calculated emission ratios compare to literature reported values from similar sources?

Thanks for pointing out this important detail. We now updated the results by using the benzene and toluene data during the night according to actinic flux data and compared the emission ratios to other studies from similar sources. As mentioned in point 12, we moved OH exposure calculation to the case studies. The text is now modified as follows:

For further insight, we focused on a time series of selected trace gases along with the correlations among them to better identify the sources of the major aliphatic carbonyls. Meanwhile, we calculated the OH exposure ($[OH]\Delta t$) based on hydrocarbon ratios (Roberts et al., 1984;de Gouw et al., 2005;Yuan et al., 2012) for these polluted regions to better understand the photochemical aging of the major carbonyls using the following equation:

$$[OH]\Delta t = \frac{1}{k_X - k_Y} \cdot \left(In \frac{[X]}{[Y]}\Big|_{t=0} - In \frac{[X]}{[Y]} \right), \qquad \text{Eq. (1)}$$

where X and Y refer to two hydrocarbon compounds with different rates of reaction with the OH radical (k). For this study, we chose toluene ($k_{OH+toluene}$: 5.63E-12 cm^3 $molecule^{-1}s^{-1}$) and benzene ($k_{OH+benzene}$: 1.22E-12 cm^3 $molecule^{-1}s^{-1}$) (Atkinson and Arey, 2003), because both compounds were measured by PTR-ToF-MS at high frequency and these values showed a good agreement with values measured by GC-FID (Figure S2). The approach detailed by Yuan et al. (2012) was applied to determine the initial emission ratio $\frac{[X]}{[Y]}\Big|_{t=0}$ in those two regions by only including nighttime data of benzene and toluene. We obtained initial emission ratios (toluene to benzene ratio) of 1.38 for the Arabian Gulf and 2.12 for the Suez region. Koss et al. (2017) summarized the toluene to benzene ratios observed in various locations and showed that urban and vehicle sources tend to have higher toluene to benzene ratio (mean ~ 2.5) than the ratios of oil & gas sources (mean ~ 1.2). Therefore, the toluene to benzene ratios obtained for those two regions agreed well with other studies done with similar emissions sources. The corresponding correlation plots of toluene and benzene for those two regions can be found in Figure S3."

15) **L329, "In general, the direct oxidation fraction varied from area to area for C1 to C3 carbonyls (formaldehyde, acetaldehyde and acetone)." Give the range of fractions for each species.**

We removed the part regarding the empirical calculations and corrected relevant text in the revised manuscript.

16) **L414, Section 3.3 title "air chemistry of unsaturated carbonyls" does not reflect the discussion in this section, which is more focused on potential precursor/source of large unsaturated carbonyls.**

We corrected the title to "Potential precursors and sources of unsaturated carbonyls" in the revised manuscript.

17) **L470-471. "A strong natural non methane hydrocarbon source from deep water in the Northern Red Sea was implemented in the model (Bourtsoukidis et al., 2020)." This needs to be added in the method section where you describe the model configuration. There you have stated that the configuration is following Bourtsoukidis et al. 2020, but also worth pointing out that this newly proposed source is applied.**

Thanks for the suggestion. We now add this information to the method section 2.5 as "The model configuration in the study is the same as the model applied in Bourtsoukidis et al. (2020), where an extra natural non-methane hydrocarbon source (ethane and propane) was implemented."

18) **L520. "the air-sea submodel (Pozzer et al., 2006) was augmented to include …" what is the air-sea submodel? Is that a model coupled into the EMAC model? Or it's a separate model you used to get the ocean acetaldehyde concentration field? Need to clarify. And if it's coupled to EMAC, need to add in the EMAC model configuration section.**

This was also mentioned by reviewer 1. The AIRSEA submodel was coupled into the EMAC. We now added this information in the method section 2.5: "The exchange of organic compounds between ocean and atmosphere was considered in EMAC via AIRSEA submodel, described in detail in Pozzer et al. (2006). The transfer velocity is calculated online and the concentration in the water is prescribed by the user. For acetone, a constant water concentration of 15 nmol/L is used, following the suggestion of Fischer et al. (2012)."

19) **P21, Section 3.5.3, "Anthropogenic primary sources". Comparing to Section 3.5.1 and 3.5.2, this section is lack of analysis-based discussion.**

Also, if the revised version kept the same length or longer, better to separate into two paragraphs, with the first one discussing the underestimated anthropogenic primary emissions of acetaldehyde precursors (L550-558) and the second talking about underestimated anthropogenic primary emissions of acetaldehyde itself (L558-end).

We reorganized this section as suggested in the revised manuscript. In terms of the analysis-based discussion, we included one more plot in the supporting information (Figure S9. Time series of ozone mixing ratios and measurement to model ratios of acetaldehyde, propene and ethene over the Arabian Gulf) inspired by the referee's later comments to further address the potential contribution to carbonyls from nighttime ozonolysis of alkenes.

L551. Does the carbon number of unmeasured hydrocarbons start from C12? Figure 3 and 6 only considered HCs up to C8.

Thanks for spotting this mistake. We corrected C12 to C8.

L552-555. "Bourtsoukidis et al. (2020) compared measured hydrocarbons (ethane, propane, ethene etc.) with the results from model simulations (the same model used in this study) and periodically found significant model underestimation in both regions. This indicates that not all sources were present in the model's emission inventory". This seems to argue that the low bias of ethane, propane, ethene, etc. in the model inventory is one reason for the model underestimation of acetaldehyde, which contradicts the argument (P13, L340-349) based on Figure 3. The argument of non-measured and non-modeled HCs before this sentence is fine, but this sentence makes it unclear.

As the Figure 3 related to empirical calculations of carbonyls was removed from the revised manuscript, there should be no confusion. We modified the sentences to make it clearer to demonstrate the contribution from underestimation of precursors in the revised manuscript: "Bourtsoukidis et al. (2020) compared measured hydrocarbons (ethane, propane, and butane) with the results from model simulations (the same model used in this study with deep water source implemented). The model was able to mostly reproduce the measurement over different regions expect for periodically significant model underestimations in Suez and Arabian Gulf, in which local and small-scale emissions were difficult for model to capture. Therefore, an underestimation of the precursor hydrocarbons as well as those large alkanes, alkenes and cyclic hydrocarbons which were not measured (> C8) or included in the model (> C5) could be a reason for the model underestimation of acetaldehyde especially in polluted regions."

L555-550. 1) need to say that the high ozone in case studies was observed during "nighttime". 2) Does the modeled occasional underestimate of ethene persist in nighttime too?

1) We corrected the sentence to "As mentioned in the previous case studies, high ozone mixing ratios were observed over the Arabian Gulf and Suez especially during the nighttime."
2) Thanks for mentioning the alkene model comparison. We now made a new plot included in the Supporting Information showing the time series of measured to model ratios of alkenes together with ozone. Relevant text was also added in the revised manuscript: "Ethene and propene were found to be significantly underestimated during the nighttime high ozone period by a factor over 10 (Figure S9), which indicates that the nighttime ozonolysis of alkenes could be another important source for acetaldehyde, formaldehyde and other carbonyls (Atkinson et al., 1995;Altshuller, 1993) in the Arabian Gulf."

L565-567. "Therefore, the active petroleum industry located in the Arabian Gulf and intensive marine transportation in Suez are likely primary sources of acetaldehyde and other carbonyls which were not well constrained in the model." This is not convincing.

We have rewritten the sentence to raise the point without indicating certainty as follows: A possible explanation for the measurement-model discrepancy is that the active petroleum industry located in the Arabian Gulf and intensive marine transportation in Suez are primary sources of acetaldehyde and other carbonyls which were not well constrained in the model.

20) L587, "Another acetaldehyde formation pathway …" suggest starting a new paragraph to discuss pyruvic acid.

We started a new paragraph to discuss pyruvic acid in the revised manuscript.

L590-592. "Zhou and Mopper (1997) pointed out that the net exchange direction for pyruvic acid is expected to be from the air to the sea due to its high partition coefficient (high solubility). Therefore, only low levels of pyruvic acid would be expected in the remote marine boundary layer". High air-to-sea partition tendency does not guarantee low level of pyruvic acid in the air-side, if gas-liquid partitioning is the only limiting

process, its concentration in the air also depends on its degree of saturation in the surface seawater. A review of surface seawater concentration of pyruvic acid and an air-sea equilibrium calculation might be helpful.

L597-598. "As the air-sea exchange of pyruvic acid is limited, low levels of pyruvic acid were expected" Not clear what this means. How did you know that the air-sea exchange of pyruvic acid is limited? And why would that result in low levels of pyruvic acid?

As the two comments above are related, we respond to them together as follows. Regarding the potential contribution of pyruvic acid photolysis to acetaldehyde, we now adopt the reviewers suggestion to make this point clear in the revised manuscript as follows: "…Pyruvic acid has been also observed in seawater (Kieber and Mopper, 1987;Zhou and Mopper, 1997) and was found up to 50 nM in the surface water of eastern pacific Ocean (Steinberg and Bada, 1984), while acetaldehyde was not the major product of aqueous-phase photolysis of pyruvic acid (Griffith et al., 2013). Zhou and Mopper (1997) pointed out that the net exchange direction for pyruvic acid is expected to be from the air to the sea due to high solubility, with Henry's law constant of 3.1×10^3 mol m^{-3} Pa^{-1} (Sander, 2015). Moreover, partitioning to aerosols could be an important sink for pyruvic acid (Reed et al.,2014; Griffith et al., 2013) : an increasing concentration trend of pyruvic acid was observed in marine aerosols over western North Pacific Ocean (Boreddy et al., 2017). Therefore, due to limited terrestrial biogenic sources of pyruvic acid for AQABA campaign, gas-phase level of pyruvic acid was expected to be low. Limited studies reported pyruvic acid level in marine boundary layer, Baboukas et al. (2000) measured 1.1 ± 1.0 ppt of pyruvic acid above the Atlantic Ocean….."

New literature citation:

Steinberg, S. M., and Bada, J. L.: Oxalic, glyoxalic and pyruvic acids in eastern Pacific Ocean waters, Journal of Marine Research, 42, 697-708, 10.1357/002224084788506068, 1984.

Sander, R.: Compilation of Henry's law constants (version 4.0) for water as solvent, Atmospheric Chemistry and Physics, 15, 4399-4981, 10.5194/acp-15-4399-2015, 2015.

Reed Harris, A. E., Ervens, B., Shoemaker, R. K., Kroll, J. A., Rapf, R. J., Griffith, E. C., Monod, A., and Vaida, V.: Photochemical kinetics of pyruvic acid in aqueous solution, J Phys Chem A, 118, 8505-8516, 10.1021/jp502186q, 2014.

L598-599. "Even if we fully assign the m/z 89.024 to pyruvic acid, the contribution to acetaldehyde via photolysis of pyruvic acid is negligible compared other sources" Clarify why it is negligible, using 58 ppt and a 100% yield? Is this contribution detectable by instrument?

We added detailed information to clarify this point in the revised manuscript as follows: "For the AQABA PTR-ToF-MS data set, enhanced signals were observed at m/z 89.024 with the mean mixing ratio of 35-110 ppt over different regions (Table S4), which is much more abundant than reported pyruvic acid levels by Baboukas et al. (2000). This might be due to the uncertainty associated with the theoretical methods of quantification used here or the presence of isomeric compounds on that mass, since pyruvic acid was not calibrated with the standard. Even if we assume the m/z 89.024 to be entirely pyruvic acid, with 60% yield of acetaldehyde via photolysis (IUPAC, 2019), it gave maximum 13 ppt of acetaldehyde over Arabian Gulf, 5-9 ppt over other regions, which were only 0.8% - 6% of the mean mixing ratios (Table S4). Detailed information of the calculation can be found in the Supporting Information. Therefore, we conclude that the contribution from the photolysis of pyruvic acid is not an important source for the unattributed acetaldehyde during the AQABA campaign."

New citation:

IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation, (http://iupac.pole-ether.fr).

Accordingly, we added the detailed information of the acetaldehyde calculation in the supporting information as follow:

Calculation of acetaldehyde yield from pyruvic acid photolysis

In order to verify the contribution from the photolysis of pyruvic acid to acetaldehyde, we calculated the expected acetaldehyde produced through pyruvic acid photolysis over different regions assuming: (1) *m/z* 89.0234 is fully assigned to pyruvic acid; (2) the loss of pyruvic acid is only through photolysis; (3) 60% is the yield of acetaldehyde via pyruvic acid photolysis recommended by IUPAC (2019); (4) the loss of acetaldehyde is only through OH oxidation. The acetaldehyde produced via pyruvic acid photolysis can be calculated using following equation (consecutive reactions):

$$[Acetaldehyde] = [Pyruvic\ acid]\frac{J_{PA}}{k_{OH}[OH]-J_{PA}}[\exp(-J_{PA}\Delta t) - \exp(-k_{OH} \times [OH]\Delta t\ )] \qquad \text{Eq. S1}$$

$[Pyruvic\ acid]$ is the mean of pyruvic acid mixing ratio in each region. J_{PA} represents the mean photolysis rate constant of pyruvic acid during the daytime (dawn to dusk) in each region calculated from the wavelength resolved actinic flux data using quantum yield of 0.2 as suggested by IUPAC (2019). The k_{OH} is the rate constant of acetaldehyde reacting with OH radical (1.5×10^{-11} cm^3 $molecule^{-1}s^{-1}$, Table S3). The $[OH]$ concentrations in each area were the mean values during the daytime obtained from the EMAC model. The maximum acetaldehyde level as well as the corresponding time (Δt) can be derived from Eq. S1 as Δt is the only variable. The results are shown in Table S4.

Table S4. Mean photolysis rate constant of pyruvic acid, OH concentrations, relative time (Δt) needed to reach the maximum acetaldehyde yield from pyruvic acid photolysis, maximum acetaldehyde and its fraction accounting the mean level over regions.

Regions	J_{PA} (s^{-1})	OH (molecules cm^{-3})	m/z 89.0234 pyruvic acid (H^+) (ppt)	Δt (h)	Acetaldehyde maximum (ppt)	Fractions (%)
MS	3.51×10^{-5}	6.52×10^{6}	39 ± 8	5.6	5.6	1.85
SC	3.44×10^{-5}	7.42×10^{6}	42 ± 9	5.2	5.3	0.85
RSN	3.52×10^{-5}	7.14×10^{6}	35 ± 14	5.2	4.7	0.92
RSS	3.00×10^{-5}	8.74×10^{6}	61 ± 15	4.9	6.2	1.98
GA	3.11×10^{-5}	7.20×10^{6}	57 ± 12	5.5	6.8	3.60
AS	2.74×10^{-5}	4.35×10^{6}	59 ± 12	7.8	9.4	5.88
GO	3.31×10^{-5}	7.89×10^{6}	65 ± 10	5.0	7.6	2.91
AG	3.29×10^{-5}	7.81×10^{6}	110 ± 53	5.1	12.9	0.75

Minor/technical comments:

L17: typo "3,8"

It is now corrected to "3.8" in the revised manuscript.

L89: section numbering error, should start with 2.1

It is now corrected to "2.1" in the revised manuscript.

L131 and Table S1: 3σ in text but 1σ in Table S1

We corrected Table S1 to 3σ.

L135: add "of" between "number" and "three"

It is now added in the revised manuscript as "… with a carbon number of three…)

L156: "interaction with ocean/land" includes "dry deposition"

We deleted “dry deposition” in the revised manuscript.

Table 1. L193, L207. “Aliphatic CCs, Aromatic CCs, Unsaturated CCS”, “CC” hasn’t been defined.

We changed “CCs” in the table to “Carbonyls” in the revised manuscript.

Table 1. L193, L207. Extend the name of “S” to “Suez”

We made this change in the revised manuscript.

L348. “section” to “sections”

The related content was deleted. Therefore, this correction was not applicable.

Figures 4, 5, and 7. Change x-axis label from “Dateandtime” to “Dateandtime (UTC)”, or clarity in the figure caption.

We made this change in the revised manuscript.

Figures 4 and 5. missing units for several y-axis, “J_NO2”, “[OH]t”, “Wind_D”

We made this change in the revised manuscript.

Figure 7a left panel. “S” to “Suez”

We made this change in the revised manuscript.

Figure S6. Only 6 regions shown in the figure.

We added the other two regions in the Figure S6 in the revised manuscript.

Measurements of carbonyl compounds around the Arabian Peninsula: overview and model comparison ~~indicate large missing sources of acetaldehyde~~

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Abstract

Volatile organic compounds (VOCs) were measured around the Arabian Peninsula using a research vessel during the AQABA campaign (Air Quality and Climate Change in the Arabian Basin) from June to August 2017. In this study we examine carbonyl compounds ~~(CxHyO)~~, measured by a proton transfer reaction mass spectrometer (PTR-ToF-MS), and present both a regional concentration distribution and a budget assessment for these key atmospheric species. Among the aliphatic carbonyls, acetone had the highest mixing ratios in most of the regions traversed, varying from 0.43 ppb over the Arabian Sea to 4.5 ppb over the Arabian Gulf, followed by formaldehyde (measured by Hantzsch monitor, 0.82 ppb over the Arabian Sea and 3~~,8~~.8 ppb over the Arabian Gulf) and acetaldehyde (0.16 ppb over the Arabian Sea and 1.7 ppb over the Arabian Gulf). Unsaturated carbonyls (C4-C9) varied from 10 to 700 ppt during the campaign, and followed similar regional mixing ratio dependence as aliphatic carbonyls, which were identified as oxidation products of cycloalkanes over polluted areas. We compared the measurements of acetaldehyde, acetone and methyl ethyl ketone to global chemistry-transport model (EMAC) results. ~~An empirical method based on hydrocarbon ratios was applied to investigate the photochemical source strength of the aliphatic carbonyls. While the distribution and relative concentration enhancements of the C3-C8 aliphatic carbonyls could be explained by this method, that of acetaldehyde could not.~~ A ~~smaller but still~~ significant discrepancy was found for ~~when comparing measurements to global chemistry-transport model (EMAC) results,~~acetaldehyde, with the model underestimating the measured acetaldehyde mixing ratio by up to an order of magnitude. Implementing a photolytically driven marine source of acetaldehyde significantly improved the agreement between measurements and model, particularly over the remote regions (e.g. Arabian Sea). However, the newly introduced acetaldehyde source was still insufficient to describe the observations over the most polluted regions (Arabian Gulf and Suez), where model underestimation of primary emissions and biomass burning events are possible reasons.

1 Introduction

Carbonyl compounds (aldehydes and ketones) can be released into the air directly from a variety of primary biogenic and anthropogenic sources. These include biomass burning (Holzinger et al., 1999;Holzinger et al., 2005;Koss et al., 2018), fossil fuel combustion (Reda et al., 2014;Huang et al., 2018) including vehicles (Erickson et al., 2014;Dong et al., 2014), industrial solvent use (Kim et al., 2008), and natural sources including plants and plankton (Zhou and Mopper, 1997;Warneke et al., 1999; Jacob et al., 2002;Fall, 2003;Williams et. al., 2004; Bourtsoukidis et al., 2014). However, secondary production via the atmospheric oxidation of hydrocarbons is considered to be more important for many carbonyl compounds including acetone and acetaldehyde (Jacob et al., 2002; Millet et al., 2010).

Carbonyls have several important roles in the atmosphere. They form as stable intermediates directly after hydrocarbon oxidation by ~~OH~~hydroxyl radicals, O_3 or NO_3, when the peroxy radicals initially formed react with each other (permutation reactions) or with NO. Their production is linked to tropospheric ozone formation (Carlier et al., 1986) and their loss, through oxidation and photolysis, is an important source of hydroxyl and hydroperoxyl radicals ~~free radicals~~ (HO_x) in the dry upper troposphere (Colomb et al., 2006). Carbonyls serve as precursors of peroxyacetyl nitrates (PANs) which are important atmospheric NO_x (NO and NO_2~~NO_x~~) reservoir species (Finlayson-Pitts and Pitts, 1997;Edwards et al., 2014;Williams et al., 2000). Carbonyl compounds are also important for the growth of atmospheric particles (Kroll et al. 2005) thereby indirectly impacting the Earth's radiative balance. The atmospheric lifetimes of carbonyl compounds varies considerably, from less than one day for acetaldehyde (Millet et al., 2010) to more than 15 days for acetone (Jacob et al., 2002;Khan et al., 2015) in terms of tropospheric mean lifetime. A multiday lifetime means that carbonyl compounds can impact the air chemistry on local, regional and even hemispheric scales. The numerous primary and secondary sources of carbonyl compounds, as well as their multiple loss routes (photolysis, OH, NO_3 and O_3 oxidation) makes budget assessments difficult.

The most predominant atmospheric carbonyl compounds besides formaldehyde are acetaldehyde and acetone. They have been reported to vary from ~~a few hundred~~ tens or hundreds of ppt in remote areas (Warneke and de Gouw, 2001; Wisthaler, 2002; Lewis et al., 2005;White et al., 2008;Colomb et al., 2009;Read et al., 2012;Sjostedt et al., 2012;Tanimoto et al., 2014; Yang et al., 2014;Hornbrook et al., 2016; Wang et al., 2019) to several ppb in urban and polluted areas (Dolgorouky et al., 2012;Guo et al., 2013;Stoeckenius and McNally, 2014; Koss et al., 2015; Sahu et al., 2017;Sheng et al., 2018). Generally, secondary photochemical formation from various precursors is the main source ~~in those regions~~for those carbonyl compounds. However, several recent studies have shown that acetaldehyde mixing ratios in both the remote marine boundary layer and the free troposphere could not be explained by known photochemistry as implemented in various atmospheric chemistry models, which consistently underestimated the measurements by an order of magnitude or more (Singh et al., 2003;Read et al., 2012;Wang et al., 2019). Several potential additional acetaldehyde sources have been proposed including new hydrocarbon oxidation mechanisms, aerosol related sources and oceanic sources. One possible source of acetaldehyde in the remote marine boundary layer is oceanic emission from the photo degradation of colored dissolved organic matter (CDOM) in sea-surface water, where acetaldehyde could be produced together with other low molecular weight carbonyl compounds (Kieber et al., 1990;Zhou and Mopper, 1997;Sinha et al., 2007;Dixon et al., 2013). Nevertheless, due to both limited airborne and seawater measurements of acetaldehyde, the importance of oceanic emission is still under debate (Millet et al., 2010;Wang et al., 2019). In order to better understand the atmospheric budgets of acetaldehyde (and the other carbonyl compounds), it is informative to analyze a dataset of multiple carbonyl compounds in both polluted and clean environments, with influence from marine emissions, varying particulate loadings, and high rates of oxidation as shown in Figure 1, which demonstrates the main formation pathways of acetaldehyde during this campaign.

During the shipborne research campaign AQABA, carbonyl compounds were continuously measured by PTR-ToF-MS onboard a research vessel that circumnavigated the Arabian Peninsula. During the campaign, chemically distinct air masses were sampled, which had been influenced by primary emissions of hydrocarbons and inorganic pollutants (NO_x, SO_2 and CO) from petroleum industries and marine transport (Bourtsoukidis et al., 2019;Celik et al., 2019), by pollution from urban areas (Pfannerstill et al., 2019), and clean marine influenced air (Edtbauer et al., 2020). It is a unique dataset of carbonyl compounds encompassing starkly different environmental conditions from a region with few (or none) available in-situ measurements to date.

In this study, we provide an overview of carbonyl compound mixing ratios (aliphatic, unsaturated and aromatic) over the Mediterranean Sea, Suez, Red Sea, Arabian Sea and Arabian Gulf. Using an empirical method based on measured hydrocarbon precursors, we have analyzed the relative importance of the photochemical sources for the carbonyl compounds observed. The analysis is then extended to include sources and transport by using a global model EMAC (5th generation European Centre – Hamburg general ~~cir-culation~~ model, **E**CHAM5 coupled to the modular earth submodel system, **M**ESSy, applied to **a**tmospheric **c**hemistry). Model measurement differences are investigated in both clean and polluted regions, with particular emphasis on acetaldehyde.

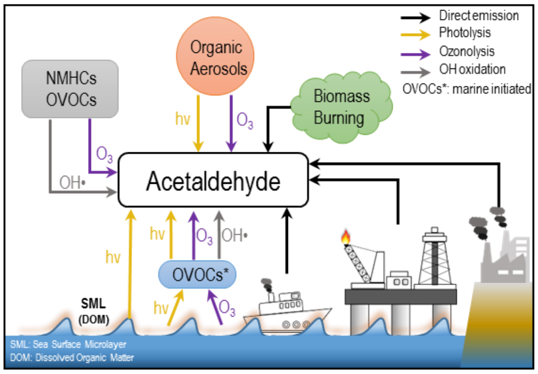

Figure 1. Diagram of possible sources and formation pathways of acetaldehyde during the AQABA campaign.

2 Methods

~~2.2~~2.1 AQABA campaign

The AQABA campaign was conducted onboard the research vessel Kommandor Iona (KI) from the end of June to the end of August 2017. The ship started from Southern France, proceeded across the Mediterranean, through the Suez Canal, around the Arabian Peninsula into the Arabian Gulf and on to Kuwait, thereafter returning along the same route. Five laboratory containers were loaded onto the vessel, containing multiple gas and particle phase measurement instruments as well as a weather station.

~~2.3~~2.2 PTR-ToF-MS

~~2.3.1~~2.2.1 Sampling and instrument set-up

A high-flow inlet (stainless steel tubing, 0.2 m diameter, 5.5 m tall and 3 m above the top of the containers and the front deck) was installed at the front of the ship where the laboratory containers were located. A high flow of air (approximately 10 m^{3}min^{-1}) was drawn through the inlet which provided a common attachment point for sub-sampling lines for all gas-phase measurement instruments. An air flow of 5 standard L min^{-1} for the first leg and 3.5 standard L min^{-1} for the second leg was pumped into the ~~VOC~~ on-board lab container through an ½" (O.D. = 1.27cm) FEP (fluorinated ethylene propylene) tubing (about 10 m long) insulated and heated to 50-60 °C. A PTFE (polytetrafluoroethylene) filter was placed at the beginning of the inlet to prevent insects, dust and particles entering the instruments. Every 2-5 days, the filter was replaced depending on the degree of pollution encountered. Inside the VOC instrument container, the PTR-ToF-MS (8000, Ionicon Analytik GmbH Innsbruck, Austria) sampled a sub-flow at 80-100 sccm through 1/8" (0.3175 cm) FEP tubing (~ 10 m in length, insulated and heated to 60 °C) from the main fast air flow and then to the instrument's PEEK (polyether ether ketone) inlet which was likewise heated to 60 °C. The inlet system was shared with total OH reactivity measurement (Pfannerstill et al., 2019).

The working principle of PTR-MS has been described in detail in previous studies (Lindinger et al., 1998;Ellis and Mayhew, 2013;Yuan et al., 2017). In brief, H_3O^+ primary ions are generated in the ion source, and then drawn into the drift tube where they interact with sampled ambient air. Inside the drift tube, VOCs with a proton affinity greater than that of H_2O (691 kJ mol^{-1}) are protonated by proton transfer from H_3O^+. The resulting secondary ions are transferred to the detector, in this case a time-of-flight mass spectrometer with mass resolution around 3500 for the first leg and 4500 for the second leg at mass 96amu. An internal standard of trichlorobenzene ($C_6H_3Cl_3$) was continuously introduced into the instrument to ensure accurate mass calibration. Every minute a spectrum with mass range (m/z) 0-450 was generated. The data reported in this study are all at 1 minute resolution unless otherwise specified.

~~2.3.2~~2.2.2 Instrument characterization

The instrument background was determined every three hours for 10 minutes with synthetic air. 4-point calibrations were performed five times during the whole campaign using a standard gas mixture (Apel-Riemer Environmental inc., Broomfield, USA) containing 14 compounds (methanol, acetonitrile, acetaldehyde, acetone, dimethyl sulfide, isoprene, methyl vinyl ketone, methacrolein, methyl ethyl ketone, benzene, toluene, xylene, 1,3,5-trimethylbenzene and α-pinene). It has been previously reported that the sensitivity of some compounds measured by PTR-MS are humidity dependent (de Gouw and Warneke, 2007). As the relative humidity (RH) was expected to be high and varying (marine boundary layer with occasional desert air influence), humidity calibration was combined with 4-point calibration by humidifying the gas mixture at different levels from 0% - 100% RH.

~~2.3.3~~2.2.3 Data analysis

The data were initially processed by the PTR Analyzer software (Müller et al., 2013) to identify and integrate the peaks. After obtaining the raw data (counts per second for each mass identified), a custom-developed python-based program was used to further process the data to final mixing ratios. For compounds present in the standard gas cylinder, interpolated sensitivities based on the five in-campaign calibrations were applied to derive the mixing ratios; while mixing ratios of the other masses were calculated by using a proton transfer reaction rate constant (k_{PTR}) of 2.0×10^{-9} cm^3 s^{-1}. The uncertainty associated with the mixing ratios of the calibrated compounds was around 6-17% (see Table S1). For the mixing ratios derived by assuming k_{PTR}, the accuracy was around ±50% (Zhao and Zhang, 2004). The detection limit (LOD) was calculated from the background measurement with 3 times the standard deviation (3σ), 52 ± 26 ppt for acetaldehyde, 22 ± 9 ppt for acetone and 9 ± 6 ppt for methyl ethyl ketone (MEK) (Table S1). Data below LOD were excluded from the data set instead of giving zero.

In this study, we have interpreted ion masses with the exact masses corresponding to $C_nH_{2n}O$, $C_nH_{2n-2}O$ and $C_nH_{2n-8}O$ as aliphatic, unsaturated and aromatic carbonyls, respectively (see exact protonated m/z in Table S2). Carbonyl compounds with a carbon number of three and above can be either aldehydes or ketones, which are not distinguishable with PTR-ToF-MS using H_3O^+ as the primary ion. However, laboratory experiments have shown that protonated aldehydic ions with carbon atoms more than three tend to lose a H_2O molecule and fragment to other masses (Buhr et al., 2002;Spanel et al., 2002). Moreover, although both ketones and aldehydes can be produced via atmospheric oxidation processes, ketones tend to have longer atmospheric lifetimes and higher photochemical yields than aldehydes as mentioned in the introduction. The ratio of measured propanal to acetone was 0.07 in the western Pacific costal region (Schlundt et al., 2017), 0.06 in an urban Los Angeles (Borbon et al., 2013) and 0.17 - 0.22 in oil &gas production regions (summarized by Koss et al., 2017). Therefore, signals on the exact mass of carbonyl compounds from the PTR-ToF-MS are expected to be dominated by ketones, particularly in regions remote from the sources.

~~2.4~~2.3 Meteorological data and other trace gases

The meteorological data were obtained by using a commercial weather-station (Sterela) which monitored wind speed, wind direction, relative humidity (RH), temperature, speed of the vessel, and GPS etc. The actinic flux was measured by a spectral radiometer (Metcon GmbH; Meusel et al., 2016). Non methane hydrocarbons (NMHC) mixing ratios were measured by a gas chromatograph with flame ionization detector (GC−FID) online with the time resolution of 50 minutes. It measured hydrocarbons (C2-C8) and aromatics (C6-C8) with the average LOD < 10 ppt for most of compounds.~~,~~ ~~for~~ For a detailed instrumental description see Bourtsoukidis et al. (2019). Formaldehyde mixing ratios were determined by a modified and optimized version of the commercially available AL4021 (Aero-Laser, Germany), which utilizes the Hantzsch technique (Stickler et al., 2006). Methane and carbon monoxide (CO) levels were monitored by a cavity ring-down spectroscopy analyzer (Picarro G2401). Ozone was measured with an absorption photometer (Model 202 Ozone Monitor, 2B Technologies, Boulder, Colorado). Due to the potential interference from sampling our own ship exhaust in which carbonyl compounds may be present (Reda et al., 2014), a filter was applied to the data set based on the wind direction and NO_x, SO_2 and ethene levels.

~~2.5~~2.4 Model simulations

The EMAC (ECHAM5/MESSy Atmospheric Chemistry) model was used to simulate atmospheric mixing ratios of several carbonyl compounds along the cruise track covered during the AQABA campaign. The EMAC model is an atmospheric chemistry-general circulation model simulating the process of tropospheric air by considering processes which could influence trace gases mixing ratios, such as transport, chemistry, interaction with ocean/land, dry deposition and so on (Pozzer et al., 2007;Pozzer et al., 2012;Lelieveld et al., 2016). The model applied in this study is a combination of the 5th generation of European Centre Hamburg general circulation model (ECHAM5) (Roeckner et al., 2006) and the 2nd version of Modular Earth Submodel System (MESSy2) (Jöckel et al., 2010), where a comprehensive chemistry mechanism MOM (Mainz Organic Mechanism) was deployed (Sander et al., 2019). The model considers direct emissions (such as anthropogenic, biogenic, biomass burning etc.), atmospheric transport and mixing, photochemical production of carbonyls (by OH, O_3 and NO_3), ~~and~~as well as physical and chemical removal processes. The global fire assimilation system was used for biomass burning emissions (Kaiser et.al., 2012). The exchange of organic compounds between ocean and atmosphere was considered in EMAC via the AIRSEA submodel, described in detail in Pozzer et al. (2006). The transfer velocity is calculated online and the concentration in the water is prescribed by the user. For acetone, a constant water concentration of 15 nmol/L is used, following the suggestion of Fischer et al. (2012). The model configuration in the study is the same as the model applied in Bourtsoukidis et al. (2020), where a natural non-methane hydrocarbon source (ethane and propane) was implemented. The model is in the resolution of T106L31 (i.e. ~ 1.1° × 1.1° horizontal resolution and , 31 vertical hybrid pressure levels up to 10 hPa) and the time resolution of 10 minutes. The measurement data of PTR-ToF-MS were averaged to 10-minute resolution to match the model data resolution for further comparison.

3 Results and discussion

Around the Arabian Peninsula, the mixing ratios of individual carbonyl compounds varied over a wide range, from tens of ppt to ppb levels. In this study, we divided the dataset geographically into eight regions (Figure 2, middle graph) to classify and characterize the primary and secondary origins of carbonyl compounds. The regional delineations were: the Mediterranean Sea (MS), Suez, Red Sea North (RSN), Red Sea South (RSS), Gulf of Aden (GA), Arabian Sea (AS), Gulf of Oman (GO) and Arabian Gulf (AG), the same as those described by Bourtsoukidis et al. (2019). Figure 2 shows the abundance of aliphatic, aromatic and unsaturated carbonyl compounds (carbonyls) for each region. Generally, aliphatic carbonyls were present at much higher mixing ratios than aromatic and unsaturated carbonyls, with smaller carbonyl compounds (formaldehyde, acetaldehyde, C3 and C4 carbonyls) dominating the distribution. The mixing ratios of aliphatic carbonyls decreased dramatically from C5 carbonyls with increasing carbon number. The box plots (Figure 2) also show that carbonyl compounds were measured at higher mixing ratios and were more variable over Suez region and the Arabian Gulf. The abundance of carbonyl compounds varied markedly from region to region with highest and lowest values found in the Arabian Gulf and the Arabian Sea, respectively. Table 1 shows the mean, standard deviation and the median values for carbonyls in each region. In the following sections, each class of carbonyl compounds are investigated in greater detail.

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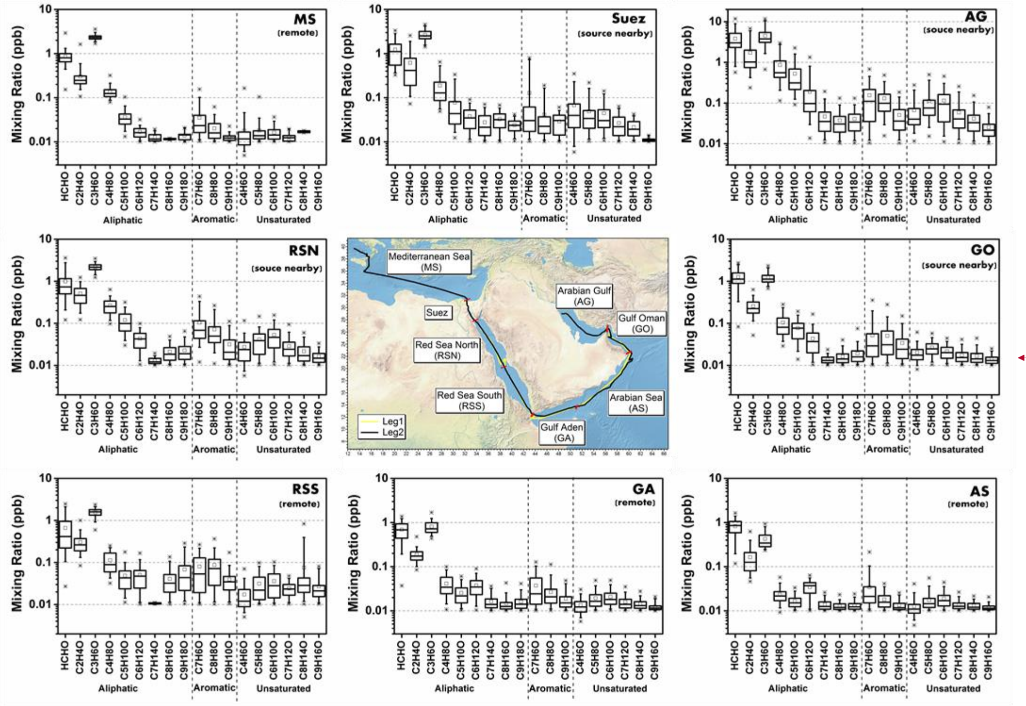

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Figure 2. Overview of mixing ratios for aliphatic, aromatic and unsaturated carbonyl compounds (CxHyO). The boxes represent 25% to 75% of the data with the central line and square indicating the median and the mean values, respectively. The whiskers show data from 5% to 95% and stars were drawn for the minimum and maximum data points within 1% to 99% of the dataset. Within brackets under the region acronyms the main characteristics of the air masses are indicated, based on non-methane hydrocarbon variability-lifetime results (b factor) from Bourtsoukidis et al. (2019)~~ and acetone mixing ratios in this study~~. The data used for map plotting was from public domain GIS data found on the Natural Earth web site (http://www.naturalearthdata.com) and was read into Igor using the IgorGIS XOP beta.

~~Table 1. Mean, standard deviation (SD) and median mixing ratios of aliphatic, unsaturated and aromatic carbonyls in different regions.~~

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~~3.1~~3.1.1 Overview~~Overview of carbonyl compounds~~

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Relatively high mean mixing ratios of aliphatic carbonyls were observed over the Arabian Gulf; the highest being acetone (C3 carbonyl compound) at 4.50 ± 2.40 ppb (median: 3.77 ppb), followed by formaldehyde at 3.83 ± 2.55 ppb (median: 3.02 ppb), acetaldehyde at 1.73 ± 1.61 ppb (median: 1.02 ppb) and MEK (C4 carbonyl compound) at 0.87 ± 0.71 ppb (median: 0.56 ppb). ~~The level of each aliphatic carbonyls over the Arabian Gulf was comparable to those previously reported for urban areas (Table 2), despite these measurements being taken at sea.~~ As the Arabian Gulf is highly impacted by the oil and gas industry, we ~~also~~ compared the measurements of the four aforementioned carbonyl compounds with those measured in the oil and gas region (Table 2). Expect for formaldehyde, acetaldehyde, acetone and MEK were lower than the mixing ratios measured in the Uintah Basin, which was influenced by intensive oil and natural gas activities (Koss et al., 2015). The general distribution of the aliphatic carbonyls in the Uintah Basin is similar to the Arabian Gulf, with acetone levels being approximately twice as those of acetaldehyde.

The carbonyl mixing ratios in the Arabian Gulf were comparable to those measured in Hickory (PA, USA) surrounded by natural [illegible] C7) measured during an aircraft measurement above the most productive oil field in the United States (Permian Basin). Within the boundary layer of the Permian Basin, C5-C7 aliphatic carbonyls had mixing ratios of 0.34 ppb, 0.08 ppb and 0.03 ppb; which are of the same magnitude but lower than the levels measured over the Arabian Gulf for C5 (0.52 ± 0.48 ppb), C6 (0.19 ± 0.25ppb) and C7 (0.05 ± 0.04 ppb) carbonyl compounds. The sources of the major carbonyls in the Arabian Gulf will be discussed in details in section 3.1.2 and 3.4.3.

In contrast, aliphatic carbonyls had much lower average mixing ratios over the Arabian Sea and the Gulf of Aden especially for C7-C9 carbonyls with mean mixing ratios below the detection limit for most of the time. During the summertime AQABA campaign, the prevailing wind direction over the Arabian Sea was southwest (Figure S1). Four-day back trajectories indicate the air was transported from the Arabian Sea (Northwestern Indian Ocean), passing East Africa coast, which brought relatively clean, photochemically aged airmasses (Bourtsoukidis et al., 2019). The mean level of acetone over the Arabian Sea (0.43 ± 0.18 ppb, median: 0.34 ppb) is close to the level measured in the marine boundary layer of Western Indian Ocean (0.49 ppb) (Warneke and de Gouw, 2001) and comparable to other reported values from ~~remote marine~~open-sea air measurement (see Table 2). -Acetaldehyde was measured at relatively low mixing ratios over the Arabian Sea (0.16 ± 0.12 ppb, median: 0.1~~2~~3 ppb), which is comparable~~lower~~ than the [illegible] MEK had slightly higher mixing ratios than those over the Arabian Sea.

The Mediterranean Sea had somewhat higher levels of aliphatic carbonyls than the clean regions (the Arabian Sea and the Gulf of Aden) but with acetone (above 2ppb) ~~still~~ dominating the distribution. Much higher acetone level than acetaldehyde level was also observed for some costal site measurement which was impacted by continental air (White et al., 2008; Schlundt et al., 2017, see [illegible] the Gulf of Oman were higher than the clean regions, while C1-C5 carbonyls were more variable over the Gulf of Oman compared to those over the Mediterranean Sea. This is probably because the Gulf of Oman connects to the Arabian Gulf where intense oil and gas industrial activities are located. Over the Gulf of Oman, polluted air from the nearby sources of the Arabian Gulf is occasionally mixed with the clean air from the open sea (the Arabian Sea) under southeast wind conditions (Figure S1).

~~4~~ Another region where abundant aliphatic carbonyls were observed was Suez region. The air in this region was mainly influenced by nearby cities and marine transportation (ship emissions within the Suez Chanel) (Bourtsoukidis et al., 2019;Pfannerstill et al., 2019). Therefore abundant precursors were available in Suez region, producing more carbonyls regionally [illegible] production contribution to the ~~acetone~~ carbonyls over Suez, the longer lived carbonyls (e.g. acetone) could be also transported from the Mediterranean Sea (where acetone was high). Four-day back trajectories indicate the air reaching Suez region was mostly originated from Europe continent passing over the Mediterranean Sea (Bourtsoukidis et al., 2019). Meanwhile, ocean uptake of acetone from the air due to polluted continental outflow (Marandino et al., 2005) as well as dilution and mixing with free tropospheric air during transport can modulate acetone mixing ratios. Although the mean mixing ratios of aliphatic carbonyls over Suez were much lower than those over the Arabian Gulf, ~~larger~~ the variations were still more significant than other regions (not including the Arabian Gulf, see Table 1)~~observed.~~.

Over the Red Sea, acetone was the most abundant aliphatic carbonyls followed by formaldehyde and acetaldehyde. The mixing ratios of acetaldehyde and acetone ~~aliphatic C2-C4 carbonyls~~ over the northern part of the Red Sea were similar to those levels measured in ~~Thompson Farm~~ western Pacific coastal regions (South China Sea, ~~a rural site in the US,~~ Table 2). It is worth noticing that the levels of aliphatic carbonyls in the northern part of the Red Sea were almost two times higher than the southern part of the Red Sea. According to the four-day back trajectories reported by Bourtsoukidis et al. (2019), the measured air masses travelled to the northern part was from southern Europe and northeast Africa while the sourthern part was more influenced by air from the northern part of the Red Sea mixed with the air masses from desertic areas of central Africa. Therefore, less primary precursors as well as carbonyls were transported to the sourhtern part of the Red Sea compared to the northern part. Moreover, the unexpected sources of hydrocarbons (ethane and propane) from Northern Red Sea deep water reported by Bourtsoukidis et al. (2020) would lead to higher carbonyl levels in the Northen part compared with the Sourthern part due to the additional precursors in the Red Sea North. However, acetaldehdye was still found to be significantly underestimated compared to the model results, even taking the deep-water source into consideration (section 3.3). This indicates [illegible]

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Table 1. Mean, standard deviation (SD) and median mixing ratios of aliphatic, unsaturated and aromatic carbonyls in different regions.

		Aliphatic Carbonyls								
		HCHO	CH3CHO	C3H6O	C4H8O	C5H10O	C6H12O	C7H14O	C8H16O	C9H18O
MS	mean	0.86	0.30	2.37	0.14	0.04	0.02	0.01	0.01	0.01
	SD	0.41	0.25	0.37	0.05	0.02	0.01	0.00	0.00	0.00
	median	0.80	0.25	2.32	0.12	0.03	0.02	0.01	0.01	0.01
Suez	mean	1.23	0.62	2.64	0.19	0.08	0.04	0.03	0.03	0.02
	SD	0.76	0.58	0.77	0.15	0.08	0.02	0.02	0.02	0.01
	median	1.11	0.42	2.52	0.13	0.04	0.04	0.02	0.03	0.02
RSN	mean	0.99	0.51	2.17	0.27	0.12	0.04	0.01	0.02	0.02

	SD	0.78	0.26	0.45	0.11	0.07	0.02	0.00	0.01	0.01
	median	0.73	0.46	2.17	0.25	0.10	0.04	0.01	0.02	0.02
RSS	mean	0.66	0.31	1.56	0.11	0.05	0.05	0.01	0.04	0.07
	SD	0.62	0.17	0.38	0.06	0.03	0.03	0.00	0.03	0.07
	median	0.40	0.26	1.60	0.09	0.04	0.05	0.01	0.03	0.04
GA	mean	0.69	0.19	0.81	0.04	0.03	0.04	0.02	0.02	0.02
	SD	0.33	0.08	0.27	0.02	0.01	0.02	0.01	0.01	0.01
	median	0.68	0.17	0.72	0.03	0.02	0.04	0.01	0.01	0.01
AS	mean	0.82	0.16	0.43	0.02	0.02	0.03	0.01	0.01	0.01
	SD	0.35	0.12	0.18	0.01	0.01	0.01	0.00	0.00	0.00
	median	0.86	0.13	0.34	0.02	0.02	0.04	0.01	0.01	0.01
GO	mean	1.27	0.26	1.33	0.10	0.08	0.04	0.01	0.02	0.02
	SD	0.59	0.12	0.40	0.06	0.04	0.03	0.00	0.01	0.01
	median	1.13	0.22	1.12	0.08	0.08	0.04	0.01	0.01	0.02
AG	mean	3.83	1.73	4.50	0.87	0.52	0.19	0.05	0.04	0.04
	SD	2.55	1.61	2.40	0.71	0.48	0.25	0.04	0.03	0.03
	median	3.02	1.02	3.77	0.56	0.31	0.10	0.04	0.03	0.03

Table 1. Continued

		Aromatic Carbonyls			Unsaturated Carbonyls					
		C7H6O	C8H8O	C9H10O	C4H6O	C5H8O	C6H10O	C7H12O	C8H14O	C9H16O
MS	mean	0.04	0.02	0.01	0.02	0.02	0.02	0.01	0.02	-
	SD	0.03	0.01	0.00	0.03	0.02	0.01	0.00	0.00	-
	median	0.02	0.02	0.01	0.01	0.01	0.01	0.01	0.02	-
Suez	mean	0.13	0.04	0.03	0.07	0.05	0.05	0.03	0.02	0.01
	SD	0.23	0.05	0.01	0.08	0.05	0.04	0.02	0.01	0.00
	median	0.03	0.02	0.03	0.04	0.03	0.03	0.02	0.02	0.01
RSN	mean	0.10	0.07	0.03	0.03	0.04	0.05	0.03	0.02	0.02
	SD	0.10	0.06	0.03	0.02	0.03	0.03	0.02	0.01	0.01

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	median	0.07	0.05	0.02	0.02	0.04	0.05	0.02	0.02	0.02
RSS	mean	0.08	0.09	0.04	0.02	0.03	0.04	0.02	0.06	0.03
	SD	0.07	0.07	0.03	0.01	0.02	0.03	0.01	0.11	0.02
	median	0.05	0.07	0.04	0.01	0.02	0.03	0.02	0.03	0.02
GA	mean	0.04	0.03	0.02	0.01	0.02	0.02	0.02	0.01	0.01
	SD	0.03	0.02	0.01	0.01	0.01	0.01	0.01	0.00	0.00
	median	0.02	0.02	0.02	0.01	0.02	0.02	0.01	0.01	0.01
AS	mean	0.03	0.02	0.01	0.01	0.02	0.02	0.01	0.01	0.01
	SD	0.04	0.01	0.00	0.01	0.01	0.01	0.00	0.00	0.00
	median	0.02	0.02	0.01	0.01	0.01	0.02	0.01	0.01	0.01
GO	mean	0.05	0.05	0.03	0.02	0.03	0.02	0.02	0.02	0.01
	SD	0.07	0.05	0.03	0.01	0.01	0.01	0.01	0.01	0.00
	median	0.03	0.03	0.02	0.02	0.03	0.02	0.02	0.01	0.01
AG	mean	0.15	0.13	0.05	0.07	0.11	0.12	0.06	0.04	0.03
	SD	0.15	0.10	0.04	0.07	0.11	0.10	0.05	0.03	0.02
	median	0.11	0.10	0.04	0.04	0.08	0.09	0.04	0.03	0.02

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Table 2. Mixing ratios (ppb) of OVOCs reported in previous observation in literature

Locations	Lon./Lat.	Height	Time	Technique	Formaldehyde	Acetaldehyde	Acetone	MEK	Literature
Open sea		m							
Tropical Atlantic Ocean	10° N-0° N 35° W-5° E	18	Oct.-Nov.	PTR-MS	n.r.	n.r.	0.53	n.r.	(Williams et al., 2004)
Atlantic Ocean	50° N-50° S 10-60° W	18	Oct.-Nov.	PTR-MS	n.r.	0.18 (Northern H) 0.08 (Southern H)	0.6 (North) 0.2 (South)	n.r.	(Yang et al., 2014)
Western North Pacific Ocean	15-20° N 137° E	6.5 -14	May	PTR-MS	n.r.	n.r.	0.20-0.70	n.r.	(Tanimoto et al., 2014)
Western Indian Ocean	12° N-5° S 43-55° E	15	Feb.-Mar.	PTR-MS	n.r.	n.r.	0.49	n.r.	(Warneke and de Gouw, 2001)
Indian Ocean	19° N-13° S 67-75° E	10	Mar.	PTR-MS	n.r.	0.32-0.42 (continental outflow) 0.18-0.21 (equatorial marine)	1.11-2.08 (continental outflow) 0.51-0.62 (equatorial marine)	n.r.	(Wisthaler, 2002)
Southern Indian Ocean	30° S-49° S 30-100° E	15	Dec.	PTR-MS	n.r.	0.12-0.52	0.42-1.08	n.r.	(Colomb et al., 2009)
Costal									
Caribbean Sea	10-30° N 60-80° W	10	Oct.	HPLC	0.61	0.57	0.40	0.03	(Zhou and Mopper, 1993)
Cape Verde Atmospheric Observatory	16.86° N 24.87° W	10	2006 - 2011	GC-FID	n.r.	0.43 (0.19-0.67)	0.55 (0.23-0.91)	n.r	(Read et al., 2012)
Appledore Island, USA	42.97° N 70.62° W	5	Jul.-Aug.	PTR-MS	n.r.	0.40	1.5	0.20	(White et al., 2008)
Mace Head, Ireland	53.3° N 9.9° W	25	Jul.-Sep.	GC-FID	n.r.	0.44 (0.12-2.12)	0.50 (0.16-1.67)	n.r.	(Lewis et al., 2005)
Canadian Archipelago	68-75° N 60-100° W	Ship cruise	Aug.-Sep.	PTR-MS	n.r.	n.r	0.34	n.r	(Sjostedt et al., 2012)
Barrow Arctic	71.30° N 156.77° W	6	Mar.-Apr.	TOGA		0.10 ± 0.20	0.90 ± 0.30	0.19 ± 0.05	(Hornbrook et al., 2016)
South China Sea, Sulu Sea	2° N-15° N 108-124° E	10	Nov.	GC-MS	n.r.	0.86	2.1	0.06	(Schlundt et al., 2017)
Oil & Gas									
Horse Pool site, Uintah Basin, USA		Ground site	2012 - 2013	PTR-MS	3.71	4.27	7.97	2.81	(Koss et al., 2015)
Central United State		<600	Mar. - April	ToF-CIMS	1.13^a	0.5	1.5	0.2	(Koss et al., 2017)
Eagle Mountain Lake site, Texas, USA		Ground site	June	PTR-MS	n.r.	n.r.	3.2 (1.2-6.7)	0.3 (0.09-0.85)	(Rutter et al., 2015)

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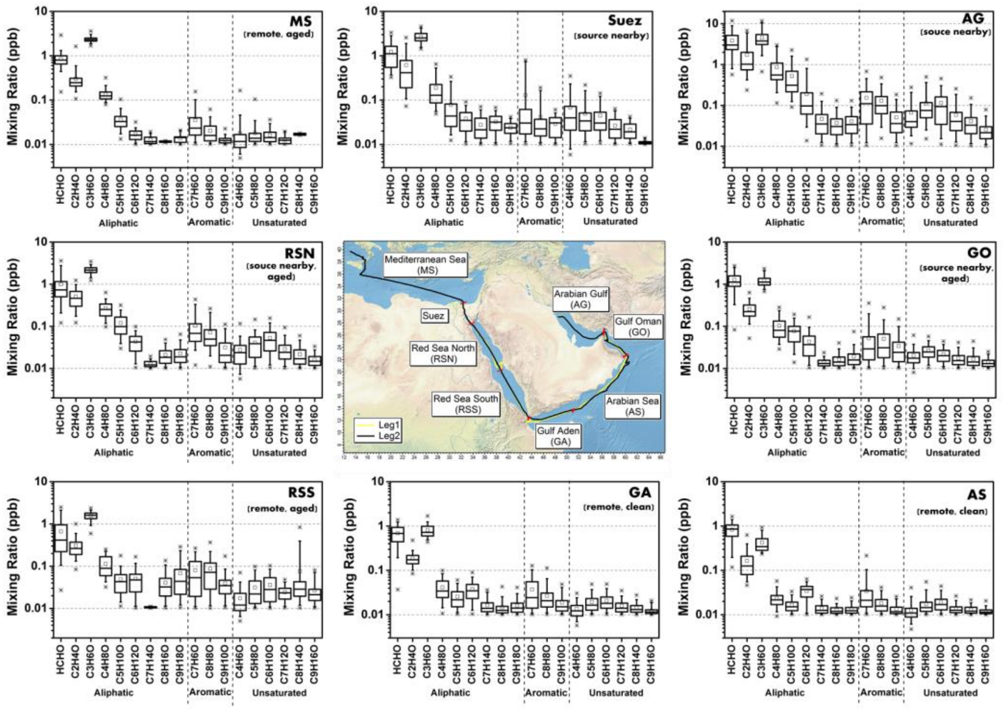
Hickory, Pennsylvania, USA	Ground site	June	PTR-MS	n.r.	1.29 (0.28-2.03)	3.22 (1.45-4.99)	0.73 (0.4-0.97)	(Swarthout et al., 2015)

n.r.: not reported in the literature.

~~5~~ a: formaldehyde was measured by laser-induced fluorescence (LIF)



~~Figure 2. Overview of mixing ratios for aliphatic, aromatic and unsaturated carbonyl compounds (CxHyO). The boxes represent 25% to 75% of the data with the central line and square indicating the median and the mean values, respectively. The whiskers show data from 5% to 95% and stars were drawn for the minimum and maximum data points within 1% to 99% of the dataset. Within brackets under the region acronyms the main characteristics of the air masses are indicated, based on variability-lifetime results (b factor) from Bourtsoukidis et al. (2019) and acetone mixing ratios in this study. The data used for map plotting was from public domain GIS data found on the Natural Earth web site (http://www.naturalearthdata.com) and was read into Igor using the IgorGIS XOP beta.~~

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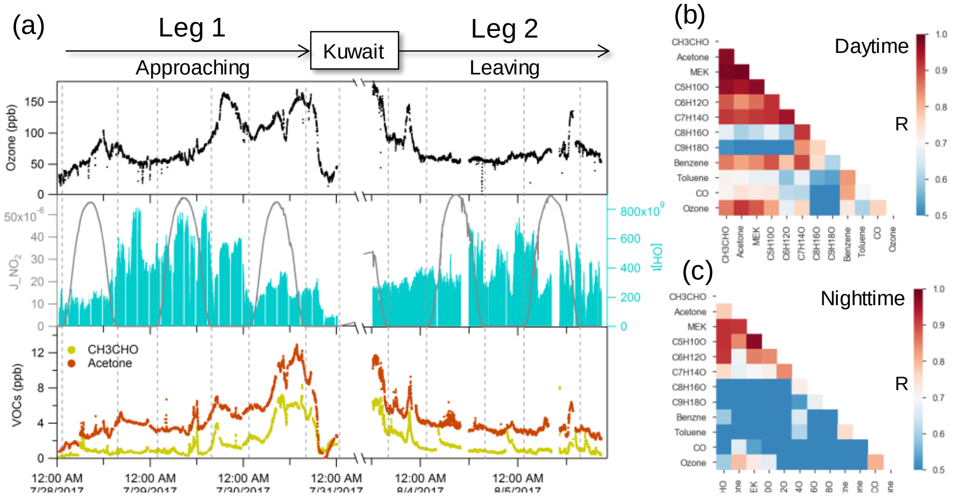
~~Table 1. Mean, standard deviation (SD) and median mixing ratios of aliphatic, unsaturated and aromatic carbonyls in different regions.~~

~~Relatively high mean mixing ratios of aliphatic carbonyls were observed over the Arabian Gulf; the highest being acetone (C3 carbonyl compound) at 4.50 ± 2.40 ppb (median: 3.77 ppb), followed by formaldehyde at 3.83 ± 2.55 ppb (median: 3.02 ppb), acetaldehyde at 1.73 ± 1.61 ppb (median: 1.02 ppb) and MEK (C4 carbonyl compound) at 0.87 ± 0.71 ppb (median: 0.56 ppb). The level of each aliphatic carbonyls over the Arabian Gulf was comparable to those previously reported for urban areas (Table 2), despite these measurements being taken at sea. As the Arabian Gulf is highly impacted by the oil and gas industry, we also compared the measurements of the four aforementioned carbonyl compounds with those measured in the oil and gas region of the Uinta Basin on land (Stoeckenius and McNally, 2014). Although the levels of three aliphatic carbonyls are higher in the Uinta Basin (mean levels of 8 ppb, 4ppb and 2.8 ppb for acetone, acetaldehyde and MEK, respectively), formaldehyde was much lower (1.9 ppb). The general distribution of the aliphatic carbonyls in the Uinta Basin is similar to the Arabian Gulf, with acetone levels being twice as those of acetaldehyde. Koss et al. (2017) reported the max boundary layer enhancement of carbonyl compounds (C2-C7) measured during an aircraft measurement above the most productive oil field in the United States (Permian Basin). Within the boundary layer of the Permian Basin, C5-C7 aliphatic carbonyls had mixing ratios of 0.34 ppb, 0.08 ppb and 0.03 ppb; which are of the same magnitude but lower than the levels measured over the Arabian Gulf for C5 (0.52 ± 0.48 ppb), C6 (0.19 ± 0.25ppb) and C7 (0.05 ± 0.04 ppb) carbonyl compounds.~~

$$[OH]\Delta t = \frac{1}{k_X - k_Y} \cdot \left(In \left.\frac{[X]}{[Y]}\right|_{t=0} - In \frac{[X]}{[Y]} \right), \qquad \text{Eq. (1)}$$

where X and Y refer to two hydrocarbon compounds with different rates of reaction with the OH radical (k). For this study, we chose toluene ($k_{OH+toluene}$: 5.63E-12 cm^3 $molecule^{-1}s^{-1}$) and benzene ($k_{OH+benzene}$: 1.22E-12 cm^3 $molecule^{-1}s^{-1}$) (Atkinson and Arey, 2003), because both compounds were measured by PTR-ToF-MS at high frequency and these values showed a good agreement with values measured by GC-FID (Figure S2). The approach detailed by Yuan et al. (2012) was applied to determine the initial emission ratio $\left.\frac{[X]}{[Y]}\right|_{t=0}$ in those two regions by only including nighttime data of benzene and toluene. We obtained initial emission ratios (toluene to benzene ratio) of 1.38 for the Arabian Gulf and 2.12 for the Suez region. Koss et al. (2017) summarized the toluene to benzene ratios observed in various locations and showed that urban and vehicle sources tend to have higher toluene to benzene ratio (mean ~ 2.5) than the ratios of oil & gas sources (mean ~ 1.2). Therefore, the toluene to benzene ratios obtained for those two regions agreed well with other studies done with similar emissions sources. The corresponding correlation plots of toluene and benzene for those two regions can be found in Figure S3.



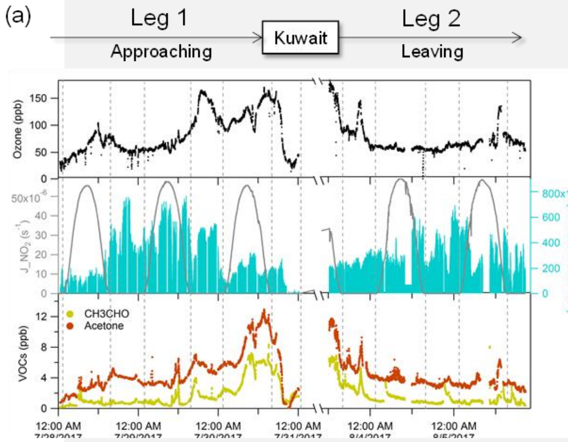
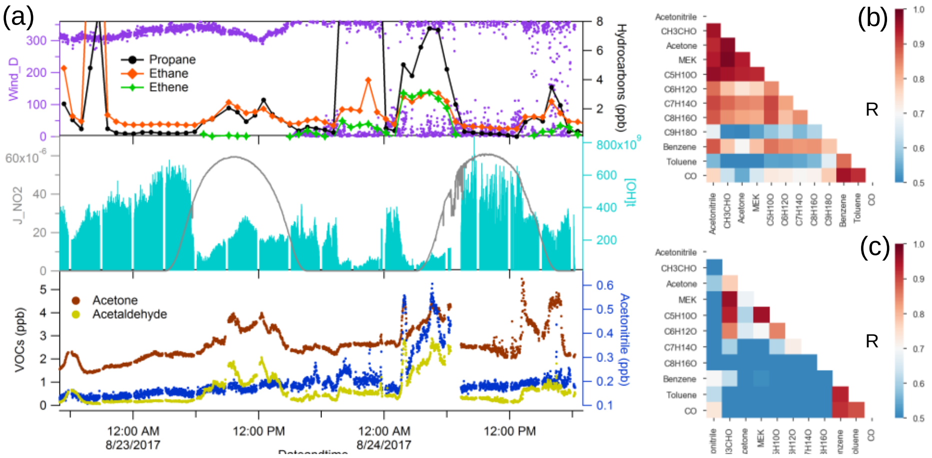


Figure ~~4~~3. Case study of the Arabian Gulf. (a) Time series of selected species measured over the Arabian Gulf; (b) daytime correlation heat map of selected species; (c) nighttime correlation heat map of selected species.

Figure ~~4~~3(a) shows the time series of acetaldehyde and acetone over the Arabian Gulf along with OH exposure ($[OH]t$) and ozone. We further separated the data into daytime and nighttime and calculated correlations among the carbonyls and other selected species (see Fig. 4b and c). Aliphatic carbonyls were well correlated with each other during the daytime and ozone had a generally good correlation with C2-C7 carbonyls (r > 0.7) during the daytime but a much lower correlation during the night, indicating ozone and carbonyls were co-produced via photochemical oxidation. Tadic et al. (2020) reported the net ozone production rate over the Arabian Gulf (32 ppb d^{-1}) was the greatest over the Arabian Peninsula. They show that strong ozone forming photochemistry occurred in this region, which would lead to abundant secondary photo-chemically produced products (including carbonyls). However, it should be noted the good correlation between ozone and carbonyls could in part be due to carbonyls co-emitted with ozone precursors (hydrocarbons) as primary emissions. ~~This further emphasizes the importance of local photochemical production of aliphatic carbonyls over the Arabian Gulf, as suggested in previous section 3.2.1. Meanwhile, a~~Is shown in Figure ~~4~~3 (a), the calculated OH exposure was high during the first night in leg 1, where an elevation of acetone mixing ratio was observed while the mixing ratio of acetaldehyde remained relatively constant. With limited OH radical abundance during the nighttime, the increased OH exposure indicates that the air reaching the ship was photochemically processed (aged). Therefore, the increase of acetone was mainly from long-distance transport as acetone has a much longer atmospheric lifetime than acetaldehyde. As the ship approached Kuwait, the calculated OH exposure was low (starting from 7/30/2017, 12:00 am UTC), which is an indicator of nearby emission sources. The lifetime of the OH radical derived from the measured OH reactivity also decreased from ~ 0.1 s to ~ 0.04 s during the same period (Pfannerstill et al., 2019). Oil fields and associated refineries are densely distributed in the northwest of the Arabian Gulf region (United States Central Intelligence Agency). The air reaching the ship when mixing ratios of acetone and acetaldehyde were highest was mainly from the Northwest (Iraq oil field region) according to the back trajectories (Bourtsoukidis et al., 2019). This suggests that the air masses encountered in Northwest Arabian Gulf were a combination of fresh emissions from nearby sources and photochemically processed air transported from elsewhere. During the second leg, relatively low mixing ratios were identified in the same region (Northwest Arabian Gulf), which was mainly due to a greater influence of air masses originating from less populated desert regions of Northeast Iran (Bourtsoukidis et al., 2019) with much less influence from the oil field emissions, meaning less precursors were available for carbonyl production. Several plumes (extending over 2-3 hours) of elevated carbonyls with increased ozone were observed during the nighttime for both legs (Fig. 4a), indicating transport of highly polluted air.



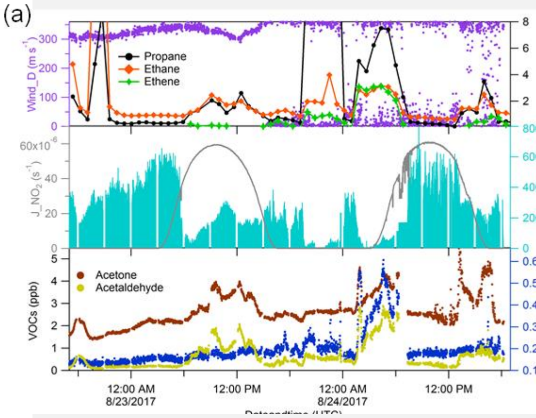


Figure ~~5~~4. Case study of Suez. (a) Time series of selected species measured over Suez; (b) correlation heat map of selected species during biomass burning plume (UTC 01:00 -06:00 August 24th 2017); (c) correlation heat map of selected species without the period of biomass burning plume.

For the Suez region (Gulf of Suez and Suez Canal), data were only available for the second leg. A significant increase of acetonitrile (over 400 ppt) was observed just before entering the Great Bitter Lake (see Figure ~~5~~4a), indicating an increasing influence of biomass burning on the air composition (Lobert et al., 1990). Carbonyl compounds are important primary emissions in fresh biomass burning plumes (Holzinger et al., 1999;Schauer et al., 2001;Holzinger et al., 2001;Koss et al., 2018) as well as being formed as secondary products in more aged plumes (Holzinger et al., 2005). We further investigated the correlation coefficient among carbonyls during the biomass burning plume (Figure ~~5~~4b) in Suez. Carbonyls had a high correlation with acetonitrile, benzene and among themselves, particularly for smaller carbonyls (acetaldehyde, C3-C5 carbonyls). The biomass burning emissions were probably transported by on the prevailing northerly wind (Figure S1) above Northeast Egypt ~~(southern side of Suez Canal)~~ where crop residues especially rice straw is often directly burned in the open fields (Abdelhady et al., 2014;Said et al., 2013;Youssef et al., 2009). Besides the direct biomass burning emission, the high mixing ratios and the good correlations of carbonyls could also have resulted from other sources as hydrocarbons (alkanes, alkenes and aromatics) which were elevated at the same time. Similar to conditions identified over the Arabian Gulf, elevated OH exposure accompanied with increasing acetone mixing ratio was observed during the first night over the Gulf of Suez, indicating aged air mass transportation. The OH exposure was then significantly lower during the daytime, when mixing ratios of carbonyls and alkanes increased as well. This indicates the presence of emission sources nearby. Oil refineries located in the coastal side of Suez and oil tank terminals located in the northern part of the Gulf of Suez are likely sources.

[illegible]

The mixing ratios of unsaturated carbonyls were generally low with values below 30 ppt over the Mediterranean Sea and the clean regions (the Arabian Sea and the Gulf of Aden, 12 - 21 ppt). The Red Sea region and the Gulf of Oman had slightly higher levels (13 – 60 ppt). The highest values were again observed in the Arabian Gulf (25 – 115 ppt) followed by Suez (11 – 68 ppt). The numbers represent the range of the mean mixing ratios of unsaturated carbonyls in each region. In terms of the mixing ratio distribution (Figure 2), the peak value was usually observed at C5 or C6 unsaturated carbonyls over most regions except for Suez where C4 carbonyl had the highest mixing ratio. Based on chemical formulas, unsaturated carbonyls can be either cyclic carbonyl compounds or carbonyls containing a carbon-carbon double bond. Therefore, the air chemistry could differ considerably depending on the compound assignment. A detailed analysis of the chemistry of the unsaturated carbonyls measured will be given in the following section 3.2.2.

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Regional variability was also observed for aromatic carbonyls with highest levels observed over the Arabian Gulf and Suez, and much lower mixing ratios over the Arabian Sea, Mediterranean Sea and Gulf of Aden (Table 1). Several studies using PTR-MS have reported values for m/z 107.049 (C7 aromatic carbonyls) attributed to benzaldehyde (Brilli et al., 2014; Koss et al., 2017;Koss et al., 2018), m/z 121.065 (C8 aromatic carbonyls) attributed to tolualdehyde (Koss et al., 2018) or acetophenone (Brilli et al., 2014) and m/z 135.080 (C9 aromatic carbonyls) attributed to methyl acetophenone (Koss et al., 2018) or benzyl methyl ketone (Brilli et al., 2014) or 3,5-dimethylbenzaldehyde (Müller et al., 2012). Atmospheric aromatic carbonyls are produced via photochemical oxidation of aromatic hydrocarbons (Finlayson-Pitts and Pitts Jr, 1999;Wyche et al., 2009;Müller et al., 2012) and benzaldehyde was reported as having primary sources from biomass burning and anthropogenic emissions (Cabrera-Perez et al., 2016). Around the Arabian Peninsula, the level of aromatic carbonyls declined with increasing carbon number over most of the regions except in the Red Sea South where C8 carbonyls were slightly higher than C7 (Figure 2). Interestingly, only in the Suez region, were the C7 aromatic carbonyls more abundant than other aromatic carbonyls, whereby the mean value (128 ± 229 ppt) was much higher than the median value (30 ppt), indicating strong primary sources of benzaldehyde in Suez. Otherwise, toluene was found to be more abundant over Suez with mean mixing ratios of 271 ± 459 ppt than over other regions (the mean over the Arabian Gulf: 130 ± 160 ppt) which would also lead to higher benzaldehyde as it is one of the OH-induced oxidation products of toluene via H-abstraction (Ji et al., 2017).

~~5.2.2~~3.2.2 Potential precursors and sources of unsaturated carbonyls

~~The primary emission sources in the Arabian Gulf and Suez regions are quite different. While the Arabian Gulf is dominated by oil and gas operations, Suez is more influenced by ship emissions and urban areas (Bourtsoukidis et al., 2019). Carbonyl compounds were most abundant in these two areas. As mentioned before, photochemical oxidation contributed a large fraction to acetone and the larger aliphatic carbonyls over the Arabian Gulf and Suez areas, but could not explain the high level of acetaldehyde measured in both regions. For further insight, we focused on a time series of selected trace-gases along with the correlations among them to better identify the sources of the major aliphatic carbonyls.~~ ketone and methacrolein (C4 carbonyls) which are frequently reported as the oxidation products of isoprene (Williams et al., 2001; Fan and Zhang, 2004;Wennberg et al., 2018). According to the GC-FID measurement, isoprene was below the detection limit for most of the time during the AQABA cruise with the highest values observed in Suez (10 - 350 ppt). This shows that the AQABA campaign was little influenced by either terrestrial or marine isoprene emissions. However, we observed unexpected high levels on mass 69.070, which is usually interpreted as isoprene for PTR-MS measurements. Significant enhancements were even identified while sampling our own ship exhaust (in PTR-MS but not GC-FID), suggesting the presence of an anthropogenic interference at that mass under these extremely polluted conditions. Several studies have reported possible fragmentations of cyclic alkanes giving mass (m/z) 69.070. These include: a laboratory study on gasoline hydrocarbon measurements by PTR-MS (Gueneron et al., 2015), a GC-PTR-MS study of an oil spill site combined with analysis of crude oil samples (Yuan et al., 2014) and an inter-comparison of PTR-MS and GC in an O&G industrial site (Warneke et al., 2014). From those studies, other fragmentations from C5-C9 cycloalkanes including m/z 43, m/z 57, m/z 83, m/z 111 and m/z 125 were identified together with m/z 69. Cyclic alkanes were directly measured in oil and gas fields (Simpson et al., 2010;Gilman et al., 2013; Li et al., 2017;Aklilu et al., 2018), vehicle exhaust (Gentner et al., 2012;Erickson et al., 2014), vessel exhaust (Xiao et al., 2018), accounting for a non-negligible amount of the total VOCs mass depending on the fuel type. Koss et al. (2017) reported enhancement of cyclic alkane fragment signals and increased levels of unsaturated carbonyls measured by PTR-ToF-MS over O&G region in the US. The unsaturated carbonyls (C5-C9) were assigned as oxidation products of cycloalkanes. Therefore, we examined the correlations between m/z 69.070 and other cycloalkane fragments over the Arabian Gulf and Suez, where anthropogenic primary emissions were significant. As shown in Figure ~~6~~5, m/z 83 was the most abundant fragment and it correlated better with m/z 69 than the other two masses, strongly supporting the presence of C6 cycloalkanes (methylcyclopentane and cyclohexane). The other two masses are distributed in two or three clusters, suggesting compositions of different cycloalkanes. M/z 43 and m/z 57 (fragments of C5 cycloalkanes) had lower correlations with other fragments (not shown in the graph) as they are also fragments of other higher hydrocarbons. Thereby we could assign those unsaturated carbonyls as photochemical oxidation products (i.e. cyclic ketones or aldehydes) from their precursor cycloalkanes.

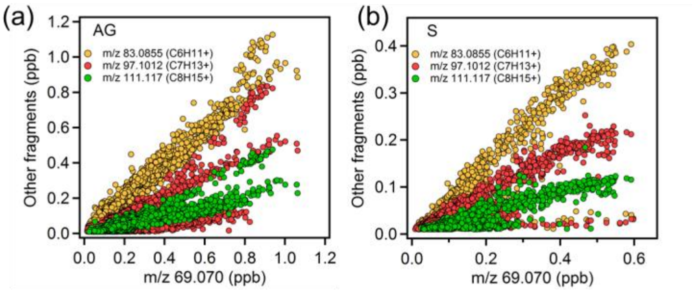


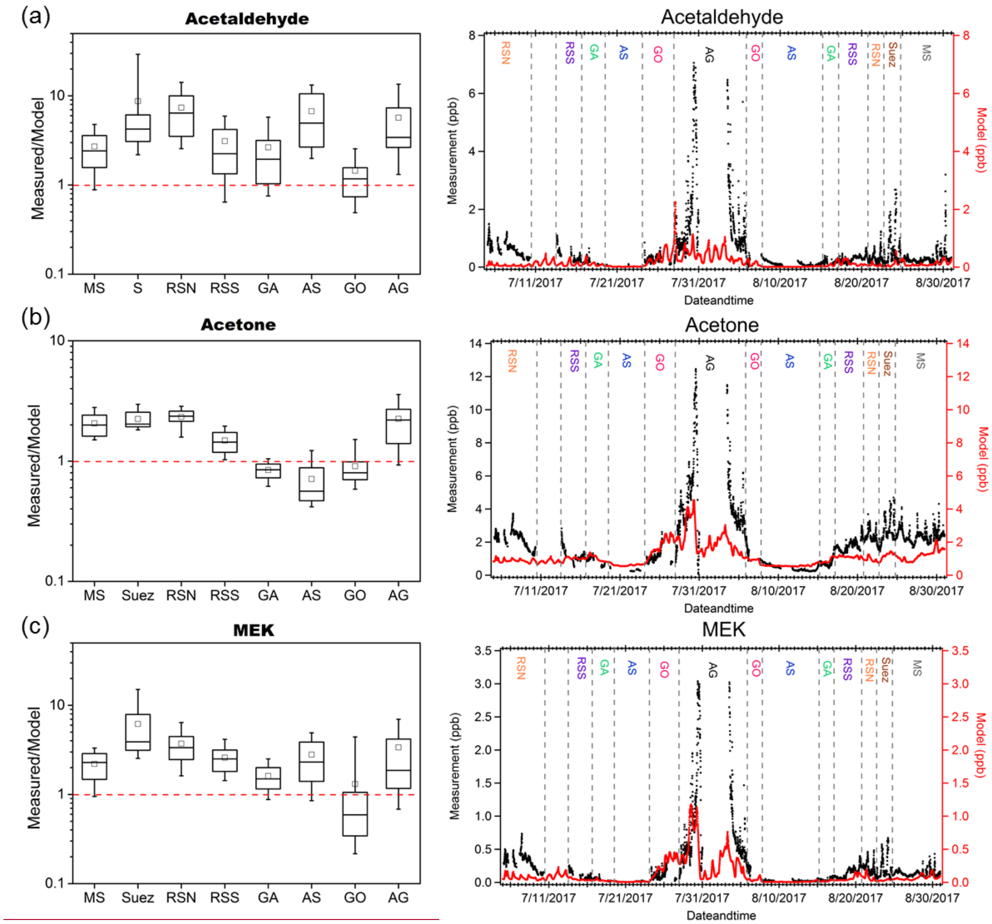
Figure ~~6~~5. Scatter plots of m/z 69.070 and other cycloalkane fragment masses over the (a) Arabian Gulf and (b) Suez region.

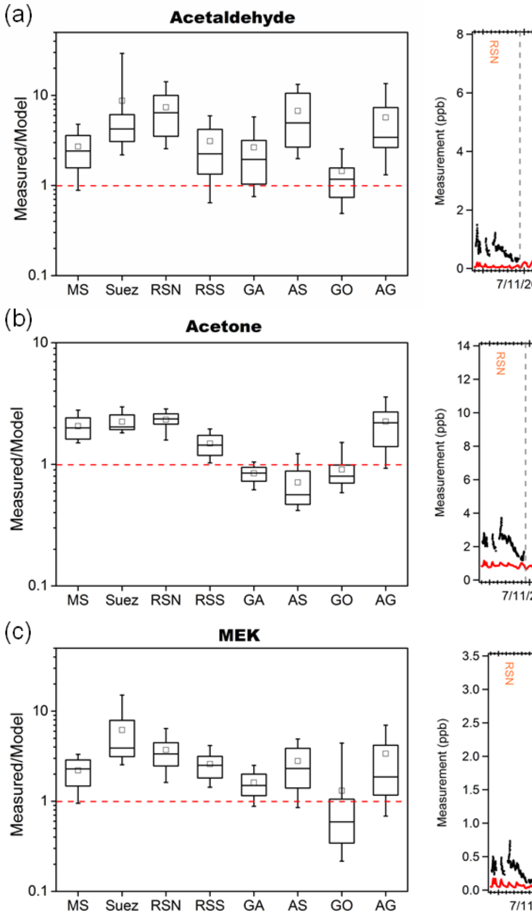
As shown in Figure 2 and Table 1, C6 unsaturated carbonyls displayed higher mixing ratios than any other unsaturated carbonyls over the Arabian Gulf while C5 unsaturated carbonyl was slightly higher than C6 in Suez. Bourtsoukidis et al. (2019) derived enhancement ratio slopes from pentane isomers and established that the Arabian Gulf is dominated by oil and gas operations and that Suez is more influenced by ship emissions. Therefore, as the Arabian Gulf had much more active O&G activities than Suez, our findings agree with Koss et al. (2017) who showed that C6 unsaturated carbonyls should be more abundant than C5 carbonyls since more precursors for C6 unsaturated carbonyls are emitted from active oil fields. It is worth mentioning that in Figure ~~6~~ 5 (b) one cluster at the bottom showed m/z 69.070 had no correlation with other three masses. Those points correspond to the time when the GC measured significant elevated isoprene while passing through the narrow Suez Canal where some vegetation (e.g. palms and some agriculture) was present close to shore, meaning m/z 69.070 during this period was isoprene. At the same time, m/z 71.049 (C4 unsaturated carbonyl) increased from 20 ppt to 220 ppt. Isoprene oxidation products (MVK and methacrolein) were probably the major contribution to the C4 unsaturated carbonyls in this period. This also explains why C4 carbonyl dominated the distribution of unsaturated carbonyls over Suez.

In the other regions (especially more remote areas), the cyclic alkane fragmentation masses had much lower abundance, leading to much less unsaturated carbonyls due to lack of precursors. Meanwhile, m/z 69.070 ($C_5H_8H^+$), m/z 83.086 ($C_6H_{10}H^+$) and m/z 97.101 ($C_7H_{12}H^+$) could also be fragmentations from corresponding aldehydes losing one water molecule as mentioned in section 2.3.3. Missing information of the chemical structure of unsaturated carbonyls and knowledge of their precursors, preclude detailed investigation of the sources of large unsaturated carbonyls in these areas.

~~5.4~~3.3 Model comparison of acetaldehyde, acetone and MEK

We compared our measurement results of acetaldehyde, acetone and MEK to those predicted by the global model "EMAC" (ECHAM5/MESSy2 for Atmospheric Chemistry). ~~The model considers direct emissions (such as anthropogenic, biogenic, biomass burning etc.), atmospheric transport and mixing, photochemical production of carbonyls (by OH, O_3 and NO_3), and physical and chemical removal processes.~~ From the results shown in ~~Figure 7~~Figure 6, the model predicted acetone much better than acetaldehyde and MEK. In general, the model broadly captured the major features identified during the campaign such as much higher levels of carbonyls mixing ratios over the Arabian Gulf and Suez and relatively low levels over the Arabian Sea. The mean measurements-to-model ratios indicated that acetone was overestimated by a factor within 1.5 over the Arabian Sea, Gulf of Aden and Gulf of Oman, and underestimated by a factor within 2.5 over the other regions. In contrast, the model underestimated MEK within a factor of 4 over most of the regions except for the Gulf of Oman where MEK was overestimated (median values were taken here as the mean values substantially deviated from the medians over Suez, Gulf of Oman and Arabian Gulf). The model underestimation was most significant for acetaldehyde, which is underpredicted by a factor (median values) of more than 6 over the Red Sea North, ~ 4 over the Arabian Sea and Arabian Gulf and between 1 and 4 over other regions. A strong natural non-methane hydrocarbon source from deep water in the Northern Red Sea was implemented in the model (Bourtsoukidis et al., 2020). Although the model representation of acetaldehyde and other carbonyls was clearly improved after including the deep water source of ethane and propane (Figure S4), the underestimation of acetaldehyde was still significant over the Red Sea North as shown in ~~Figure 7~~Figure 6(a), indicating further missing sources. For acetaldehyde and MEK, the discrepancy was also significant over the Arabian Sea where acetone was in contrast, overestimated. Since acetaldehyde had the biggest bias from the model prediction~~, both with our simple empirical calculation (section 3.2.1) and the global model,~~ we further investigate the possible missing sources of acetaldehyde.





~~Figure 7~~Figure 6. Measurement to model ratios (left) and time series (right) of measurements (in black) and model simulation (in red) of (a) acetaldehyde; (b) acetone; (c) MEK in each area. In each box plot, the box represents 25% to 75% of the data set with central line and square indicating the median value and the mean value respectively. The whiskers show data from 10% to 90%. The red dashed lines represent the 1:1 ratio.

~~5.5~~3.4 Missing sources of acetaldehyde

In this section we investigate the following processes as potential sources of acetaldehyde: (1) production as an inlet artifact, (2) oceanic emission of acetaldehyde, (3) anthropogenic primary sources, (4) biomass burning sources, and (5) other possible secondary formation pathways.

~~5.5.1~~3.4.1 Inlet artifact

Northway et al. (2004) and Apel et al. (2008) reported that heterogeneous reactions of unsaturated organic species with ozone on the wall of the Teflon inlet can cause artifacts signal of acetaldehyde but not to acetone. During AQABA, the highest and the most variable ozone mixing ratios were observed during the campaign over the Arabian Gulf (mean: 80 ± 34 ppb) and the Red Sea North (66 ± 12 ppb), where a modest correlation was found between acetaldehyde and ozone over the Arabian Gulf (r^2=0.54) and no significant correlation over the Red Sea North (r^2=0.40). However larger correlation coefficients were identified between ozone and other carbonyls over the Arabian Gulf (see Figure S5), which suggests that the correlation was due to atmospheric photochemical production rather than artifacts. Moreover, acetaldehyde was found to have a much worse correlation with ozone during the nighttime compared to the correlation during the daytime over the Arabian Gulf (Figure ~~4b~~ 3b and c), which also indicates that inlet generation of acetaldehyde was insignificant. Over other regions, especially the remote area (the Arabian Sea and Gulf of Aden), ozone was relatively constant and low, with poor correlation with acetaldehyde mixing ratios. Although we cannot completely exclude the possible existence of artifacts, the interference is likely to be insignificant in this dataset.

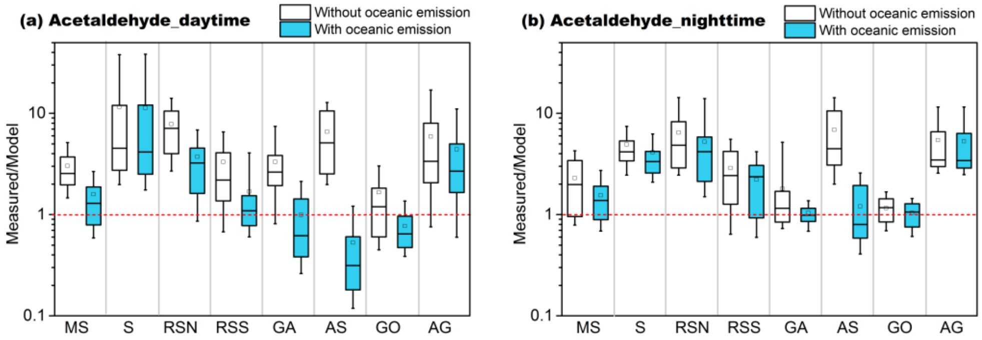
~~5.5.2~~3.4.2 Oceanic emission

A bias between measured acetaldehyde and global model simulations has been observed in previous studies conducted in the remote troposphere (Singh et al., 2003; Singh, 2004;Wang et al., 2019) and in the marine boundary layer (Read et al., 2012). The aforementioned studies emphasized the potential importance of the sea water acting as a source of acetaldehyde emission via air-sea exchange. No significant correlation was found between acetaldehyde and DMS, a marker of marine biogenic emission which is produced by phytoplankton in seawater (Bates et al., 1992) (see Figure S6). This indicates that the direct biogenic acetaldehyde emissions from the ocean are probably insufficient to explain the measured acetaldehyde~~This indicates that the source of acetaldehyde was probably not from direct biogenic production, which has been reported by Mungall et al. (2017).~~. More likely, acetaldehyde and other small carbonyl compounds can be formed in the sea especially in the surface microlayer (SML) via photodegradation of colored dissolved organic matter (CDOM) (Kieber et al., 1990;Zhou and Mopper, 1997;Ciuraru et al., 2015). Zhou and Mopper (1997) calculated the exchange direction of small carbonyls based on measurement results and identified that the net flux of acetaldehyde was from sea to the air whereas formaldehyde was taken up by the sea. Sinha et al. (2007) characterized air-sea flux of several VOCs in a mesocosm experiment and found that acetaldehyde emissions were in close correlation with light intensity (r=0.7). By using a 3-D model, Millet et al. (2010) estimated the net oceanic emission of acetaldehyde to be as high as 57 Tg a^{-1} (in a global total budget: 213 Tg a^{-1}), being the second largest global source. A similar approach was applied in a recent study done by Wang et al. (2019), reporting the upper limit of the net ocean emission of acetaldehyde to be 34 Tg a^{-1}. Yang et al. (2014) quantified the air-sea fluxes of several OVOCs over Atlantic Ocean by eddy covariance measurements, showing ocean is a net source for acetaldehyde. Although Schlundt et al. (2017) reported uptake of acetaldehyde by the ocean from measurement-inferred fluxes in western Pacific coastal regions, t~~T~~o our knowledge, there is no ~~clear~~ direct experimental evidence showing the ocean to be a sink for acetaldehyde.

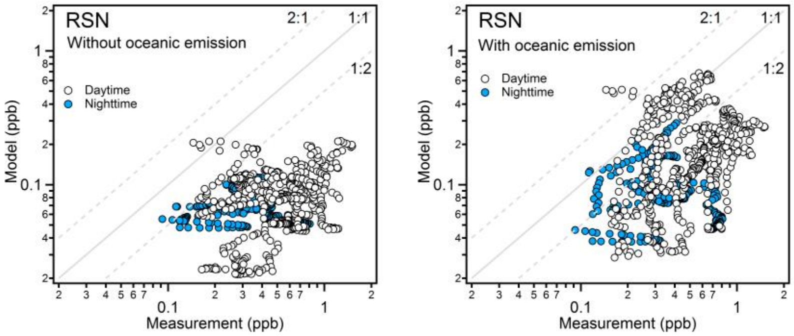
In order to test the importance of the oceanic emission of acetaldehyde, we implemented this source in EMAC model. The measured sea water concentration of acetaldehyde was not available for the water area around the Arabian Peninsula. Wang et al. (2019) estimated the global average acetaldehyde surface seawater concentrations of the ocean mixed layer using a satellite-based approach similar to Millet et al. (2010), where the model estimation agreed well with limited reported measurements. From the Wang et al. (2019) results, the averaged seawater concentration of acetaldehyde around Arabian Peninsula was generally much higher from June to August. As the photodegradation of CDOM is highly dependent on sunlight, the air-sea submodel (Pozzer et al., 2006) was augmented to include throughout the campaign a scaled acetaldehyde seawater concentration in the range of 0 ~ 50 nM according to the solar radiation (Figure S7). With this approach, the average of acetaldehyde seawater concentration estimated by the model is 13.4 nM, a reasonable level compared to predicted level by Wang et al. (2019).

After adding the oceanic source of acetaldehyde, the model estimation was significantly improved (~~Figure 8~~Figure 7). As the oceanic source in the model is scaled according to the solar radiation, the measurement-to-model ratios were more strongly reduced during the day compared to the night. With oceanic emission included, the model underestimation was less significant, within a factor of 3 during the day and 4 during the night over the Mediterranean Sea, Red Sea and Gulf of Aden. The most significant improvement was identified over the Red Sea North. As shown in ~~Figure 9~~Figure 8, the model had much better agreement with the measurement after adding the oceanic source. The scatter plots for other regions can be found in Figure S8. Over the Arabian Sea, the model significantly overestimated acetaldehyde mixing ratios, indicating the input sea water concentration of acetaldehyde might be too high. The SML layer starts to be effectively destroyed by the wave breaking when the wind speed exceeds than 8 m s^{-1} (Gantt et al., 2011). As the average wind speed over the Arabian Sea was the highest among the cruised areas (8.1 ± 2.4 m s^{-1}, Figure S1), less contribution from the CDOM photo degradation to acetaldehyde in the surface sea water would be expected. For the Suez region, due to the limited model resolution ($1.1^o \times 1.1^o$), little sea water was identified in the model, leading to negligible influence from the oceanic source.

Model underestimation of acetaldehyde especially over the Suez, Red Sea and Arabian Gulf is also likely to be related to the coarse model resolution (~ $1.1^o \times 1.1^o$) (Fischer et al., 2015). Where model grid points contain areas of land the higher and more variable terrestrial boundary layer height impacts the model prediction whereas the measurements may only by influenced by a shallower and more stable marine boundary layer.



~~Figure 8~~Figure 7. Acetaldehyde measurement to model ratios without the oceanic source (white boxes) and with the oceanic source (blue boxes) in the model during (a) daytime and (b) nighttime in different regions. The boxes represent 25% to 75% of the data set with the central line and square indicating the median and mean values, respectively. The whiskers show data from 10% to 90%. The red dashed lines represent the 1:1 ratio.



~~Figure 9~~Figure 8. Observed and simulated mixing ratios of acetaldehyde over the Red Sea North without oceanic emission (left) and with oceanic emission (right). The data points are separated into day- and nighttime according to solar radiation.

~~5.5.3~~3.4.3 Anthropogenic primary sources

Over the Arabian Gulf and Suez, the intensive photochemical production of carbonyls is apparent. Bourtsoukidis et al. (2020) compared measured hydrocarbons (ethane, propane, and butane~~ethene etc.~~) with the results from model simulations (the same model used in this study with the newly discovered deep water source implemented). The model was able to reproduce the measurement over most regions expect for some significant model underestimations in Suez and Arabian Gulf, in which local and small-scale emissions were difficult for the model to capture. Therefore, an underestimation of the precursor hydrocarbons, ~~especially~~ as well as those large alkanes, alkenes and cyclic hydrocarbons which were not measured (> ~~C12~~C8) or included in the model (> C5) could be a reason for the model underestimation of acetaldehyde especially in polluted regions ~~and other carbonyls. Bourtsoukidis et al. (2020) compared measured hydrocarbons (ethane, propane, ethene etc.) with the results from model simulations (the same model used in this study) and periodically found significant model underestimations in both regions~~. In addition, ~~. This indicates that not all sources were present in the model's emission inventory.~~ a~~A~~s mentioned in the previous case studies, high ozone mixing ratios were observed over the Arabian Gulf ~~and Suez~~especially during the nighttime. ~~With large amounts of alkenes present in those regions, which the model occasionally underestimated~~Ethene and propene were found to be significantly underestimated during the nighttime high ozone period by a factor over 10~~,~~ (Figure S9), ~~-~~which indicates that the nighttime ozonolysis of alkenes could be another important source for acetaldehyde, formaldehyde and other carbonyls (Atkinson et al., 1995;Altshuller, 1993) in the Arabian Gulf.

~~-~~Acetaldehyde, an oxygenated VOC, is not generally considered as an important primary emission from oil and gas field but instead a photochemical product of hydrocarbon oxidation (Yuan et al., 2014;Koss et al., 2015;Koss et al., 2017). In contrast, primary sources of formaldehyde from oil and gas production processes including both combustion and non-combustion process have been ascertained (Vaught, 1991). Le Baron and Stoeckenius (2015) concluded in their report of the Uinta Basin winter ozone study that besides formaldehyde, the other carbonyls were poorly understood in terms of their primary sources. Acetaldehyde and other carbonyls (aldehydes and ketones) have been reported as primary emissions from fossil fuel combustion including ship emissions (Reda et al., 2014;Xiao et al., 2018;Huang et al., 2018) and vehicle emissions (Nogueira et al., 2014;Erickson et al., 2014;Dong et al., 2014). ~~-~~A possible explanation for the measurement-model discrepancy is that the active petroleum industry located in the Arabian Gulf and intensive marine transportation in Suez are primary sources of acetaldehyde and other carbonyls which were not well constrained in the model.~~Therefore, the active petroleum industry located in the Arabian Gulf and intensive marine~~

transportation in Suez are likely primary sources of acetaldehyde and other carbonyls which were not well constrained in the model. The Suez region, where the largest acetaldehyde discrepancy was identified, had a significant influence from biomass burning (see section 3.2.2). Biomass burning emissions are notoriously difficult to model as they are highly variable both in time and space. In this study, the model failed to reproduce the acetonitrile level with a range of only 40-50 ppt rather than 100-550 ppt measured over Suez. Thus, besides the possibility of seawater emission from the Gulf of Suez and the Suez Canal, the underestimated biomass burning source in the model over Suez, will lead to an underestimation of acetaldehyde as well as other carbonyl compounds in this region.

~~5.5.4~~3.4.4 Other possible secondary formation pathways

Although the model estimation was generally improved with the addition of an oceanic source, the model to measured ratios still varied over a wide range. As mentioned above, photodegradation of CDOM on the surface of seawater is a known source for acetaldehyde although some studies focusing on real sea water samples did not observe clear diel cycles of seawater acetaldehyde (Beale et al., 2013;Yang et al., 2014). Fast microbial oxidation could be a reason (Dixon et al., 2013) while other non-light driven sources of acetaldehyde could be an alternative explanation. In a recent study, Zhou et al. (2014) reported enhanced gas-phase carbonyl compounds including acetaldehyde during a laboratory experiment of ozone reacting with SML samples, indicating acetaldehyde could also be produced under non-light driven heterogeneous oxidation. Wang et al. (2019) ventured a hypothetical source that organic aerosol can be an extra source for unattributed acetaldehyde in the free troposphere through light-driven production and ozonolysis. However, since the yield of acetaldehyde from such reactions is unknown, large uncertainties remain. Previous studies have shown that the organic matter fraction was highest in smaller sea spray aerosols and that the aerosols contain both saturated and unsaturated fatty acids originating from the seawater surface (i.e. SML) (Mochida et al., 2002;Cochran et al., 2016). Thus, for the AQABA campaign, both photodegradation and heterogeneous oxidation could occur on the surface of sea spray and pollution associated aerosols, even over remote open ocean therefore being an extra source of acetaldehyde and other carbonyl compounds.

Another acetaldehyde formation pathway reported is gas-phase photolysis of pyruvic acid (Eger et al., 2019b;Reed Harris et al., 2016), a compound mainly of biogenic origin. Pyruvic acid has been also observed in seawater (Kieber and Mopper, 1987;Zhou and Mopper, 1997~~;Tedetti et al., 2006~~) and was found up to 50 nM in the surface water of easter pacific Ocean (Steinberg and Bada, 1984), while ~~, although~~ acetaldehyde was not the major product of aqueous-phase photolysis of pyruvic acid (Griffith et al., 2013). Zhou and Mopper (1997) pointed out that the net exchange direction for pyruvic acid is expected to be from the air to the sea due to ~~its high partition coefficient (~~high solubility~~)~~, with Henry's law constant of 3.1×10^3 mol m^{-3} Pa^{-1} (Sander, 2015). Moreover, partitioning to aerosols could be an important sink for pyruvic acid (Reed et al.,2014; Griffith et al., 2013) : an increasing concentration trend of pyruvic acid was observed in marine aerosols over western North Pacific Ocean (Boreddy et al., 2017). Therefore, due to limited terrestrial biogenic sources of pyruvic acid for AQABA campaign, gas-phase level of pyruvic acid was expected to be low. Limited studies reported ~~Therefore, only low levels of pyruvic acid would be expected in the remote marine boundary layer. p~~Pyruvic acid level in marine boundary layer, Baboukas et al. (2000) measured 1.1±1.0 ppt of pyruvic acid above Atlantic Ocean. Pyruvic acid was measured by Jardine et al. (2010) using a PTR-MS at m/z 89 in a forested environment. For the AQABA PTR-ToF-MS data set, enhanced signals were observed at m/z 89.024 with the mean mixing ratio of ~~58 ± 34~~35-110 ppt over different regions (~~the boxplot can be found in Figure S9~~ Table S4), which is much more abundant than reported pyruvic acid levels by Baboukas et al. (2000) measured above Atlantic Ocean (1.1±1.0 ppt)~~(Baboukas et al., 2000)~~. This might be due to the uncertainty associated with the theoretical methods of quantification used here or the presence of isomeric compounds on that mass, since pyruvic acid was not calibrated with the standard. ~~As the air-sea exchange of pyruvic acid is limited, low levels of pyruvic acid were expected.~~ Even if we ~~fully assign~~ assume the m/z 89.024 to be entirely pyruvic acid, with 60% yield of acetaldehyde via photolysis (IUPAC, 2019), it gave maximum 13 ppt of acetaldehyde over Arabian Gulf, 5-9 ppt over other regions, which were only 0.8% - 6% of the mean mixing ratios (Table S4). Detailed information of the calculation can be found in the Supporting Information. ~~the contribution to acetaldehyde via photolysis of pyruvic acid is negligible compared to other sources.~~ Therefore, we conclude that the contribution from the photolysis of pyruvic acid is not an important source for the unattributed acetaldehyde during the AQABA campaign.

~~6~~4 Summary and Conclusion

Observations of carbonyl compounds around the Arabian Peninsula were investigated in terms of mixing ratios abundance over different areas. Aliphatic carbonyl compounds were generally more abundant than the unsaturated and aromatic carbonyl compounds, and were dominated by low-molecular-weight compounds (carbon number less than five). Aliphatic carbonyl compounds were found at the highest mixing ratios over the Arabian Gulf followed by the Suez region, while the lowest mixing ratios were observed over the Arabian Sea and the Gulf of Aden. Over the Mediterranean Sea, aliphatic carbonyls were low except for acetone that was much higher compared to the levels observed over clean remote areas (i.e. Arabian Sea). The atmospheric composition over the Red Sea showed obvious differences between the northern and the southern part, with higher mixing ratios in the north. Similar region-dependent distributions were observed for unsaturated and aromatic carbonyls. Generally, the mixing ratios of aromatic carbonyl compounds decreased as the carbon number increased. Particularly over the Suez region, benzaldehyde (C7 aromatic carbonyls) was much more abundant than other aromatic carbonyls, indicating direct sources as well as abundant oxidation precursors. For unsaturated carbonyl compounds, C5 and C6 carbonyl compounds dominated the mixing ratio distribution, while the air chemistry highly depends on the chemical structure assignment of those masses.

~~To better understand the air chemistry of aliphatic carbonyl compounds over different regions, we used an empirical method to calculate the levels of carbonyl compounds resulting from OH oxidation of precursor hydrocarbon species. The results indicate that mixing ratios of formaldehyde and C3-C8 carbonyl compounds could, to a large part, be explained by OH initiated photooxidation in each region, especially over the Arabian Gulf and Suez region. This result indicates that photooxidation is a dominant production pathway for formaldehyde and C3-C8 aliphatic carbonyl compounds in these two regions. However, acetaldehyde from hydrocarbon precursors was not sufficient to explain the high mixing ratios observed, indicating the existence of other sources and/or formation pathways.~~ Further case studies showed that the carbonyl compounds ~~produced via photooxidation~~ were highly correlated to the high ozone levels during daytime over the Arabian Gulf while the air chemistry in Suez region was strongly influenced by regional biomass burning. Due to the unexpectedly high loading of m/z 69 (usually assigned as isoprene) observed in highly polluted regions, we further identified the correlations between m/z 69 and other fragmentation masses of cycloalkanes according to previous studies conducted in oil and gas regions (Warneke et al., 2014;Yuan et al., 2014;Koss et al., 2017). The high correlations among fragments implied the existence of cycloalkanes in the polluted regions, which could be further oxidized to unsaturated carbonyl compounds (cyclic ketones or aldehydes).

As acetaldehyde was identified as having important additional sources, we further compared the measurements of major carbonyl species (acetaldehyde, acetone and MEK) with a comprehensive global atmospheric chemistry model (EMAC). Acetaldehyde was found to have the highest discrepancy between the observations and model simulations, with the simulated values to be lower up to a factor of 10. By adding an oceanic source of acetaldehyde produced via light-driven photodegradation of CDOM in the seawater, the model estimation improved significantly, especially over the Red Sea North. With the oceanic source added, modelled acetaldehyde became slightly overestimated in clean regions, suggesting that the emission rate employed represents an upper limit. The results indicate that the ocean plays an important role in the atmospheric acetaldehyde budget, under both clean and polluted conditions. The underestimated acetaldehyde in the model is significant as it will influence the atmospheric budget of e.g. PAN. As shown in Figure 1, multiple sources and formation pathways need to be considered to better understand the atmospheric budget of acetaldehyde. Additional laboratory experiments and field measurements are necessary in order to verify all possible atmospheric formation mechanisms and to improve model simulations.

Data availability.

Data will be made available via: https://edmond.mpdl.mpg.de/imeji/

Author contributions.

AE and CS performed PTR-ToF-MS measurement and preliminary data processing. NW conducted data analysis and drafted the article. AP performed EMAC model simulation. EB and LE are responsible for NMHC measurements and data. DD, BH and HF provided formaldehyde data. Ozone and actinic flux data were contributed by JS and JNC. Methane and carbon monoxide data were provided by JP. JL designed and realized the campaign. JW supervised the study. All authors contributed to editing the draft and approved the submitted version.

Competing interest.

The authors declare that they have no conflict of interest.

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