The authors thank the reviewers for their insightful comments. The response to each comment is in blue with corresponding changes in the manuscript italicized.

Anonymous Referee #1

The manuscript presents some interesting measurements of cooking SOA by TDGC/MS. The authors adopt a new method to achieve molecular speciation of the SOA, where SOA complex mixture was deconvoluted using mass spectral fragmentation patterns to extract useful information about functional groups and carbon numbers. They also derived the parameterizations for aldehyde oxidation and used the derived parameters to predict SOA mass. The model results generally well captured the amount of SOA formed and its chemical characteristics, e.g., O/C. Overall, this paper makes a valuable contribution to current cooking SOA understanding.

Q1. Hydroxyl radicals as major oxidants are produced through the ozone photolysis in the flow tube. However, ozone concentration would be high in the flow tube. What is the influence of ozone oxidation of aldehydes in the experiments?

We agree that ozone concentration would be high in the flow tube. However, we did not observe any particle formation during ozonolysis of heated cooking oil where OH radicals were not present. Since aldehydes are major contributors of gas-phase emissions from cooking oils, we believe that during photooxidation of aldehyde precursors ozone chemistry would have negligible effect on particle formation. Furthermore, gas-phase reaction rate constant of aldehydes with ozone is of the order of 10^{-17}-10^{-18} (Atkinson and Arey, 2003; Atkinson and Carter, 1984) much lower than that of hydroxyl radicals, so the timescale for ozonolysis of aldehydes would likely be longer than the residence time in the flow tube reactor.
To understand the reactivity of C=C in unsaturated aldehydes, methacrolein is used as an example calculation to provide insights into the reaction timescales as its reaction rate constant with both OH and ozone is readily available. Following calculations shows that timescale of reaction with ozone is much slower than the residence time in the flow tube suggesting that the ozonolysis of unsaturated aldehydes play a negligible role in the formation of SOA from aldehydes or cooking emissions.
The reaction rate constant of methacrolein is obtained from Atkinson and Arey, 2003.
k_{OH} is 2.9E-11 cm^3 $molec^{-1}$ s^{-1}
k_{O_3} is 1.2E-18 cm^3 $molec^{-1}$ s^{-1}.
For the lowest OH exposure, OH conc = 2.88E8 molec cm^{-3} and ozone conc is 0.5 ppm. At these conditions, methacrolein reaction timescale with OH is $1/k_{OH}*[OH]$ = ~120 s, while reaction timescale with ozone is 1129 min.
For the highest OH exposure, OH conc = 1.1E9 molec cm^{-3} and ozone conc is 12.6 ppm. At these conditions, methacrolein reaction timescale with OH is 31 s, while reaction timescale with ozone is ~45 min.
To clarify this, following text have been added to the manuscript in L91-94, as well as above sample calculation have been added to SI in Sect. 1 in L60-73:

"…The effect of ozone on the SOA formation was found to be negligible as the reaction timescales of aldehydes with ozone were calculated to be at least 100 times longer than those with OH. A

sample calculation for methacrolein reaction timescales with OH and ozone is shown in SI in Sect. 1."

Q2. What are the criteria in selecting the model compounds (Table S2) in each functional group class (e.g., different carbon number range used for different classes)?

Based on the carbon number range of VOC precursors emitted from cooking emissions, the resulting SOA products will have a similar or lower carbon # assuming oligomerization reactions are not dominant in the system. We also expect that photooxidation reactions will lead to addition of -OH and =O groups based on knowledge about gas-phase oxidation chemistry. Based on these different combinations of functional groups we selected the model compounds in Table S2 to characterize the composition of cooking SOA.

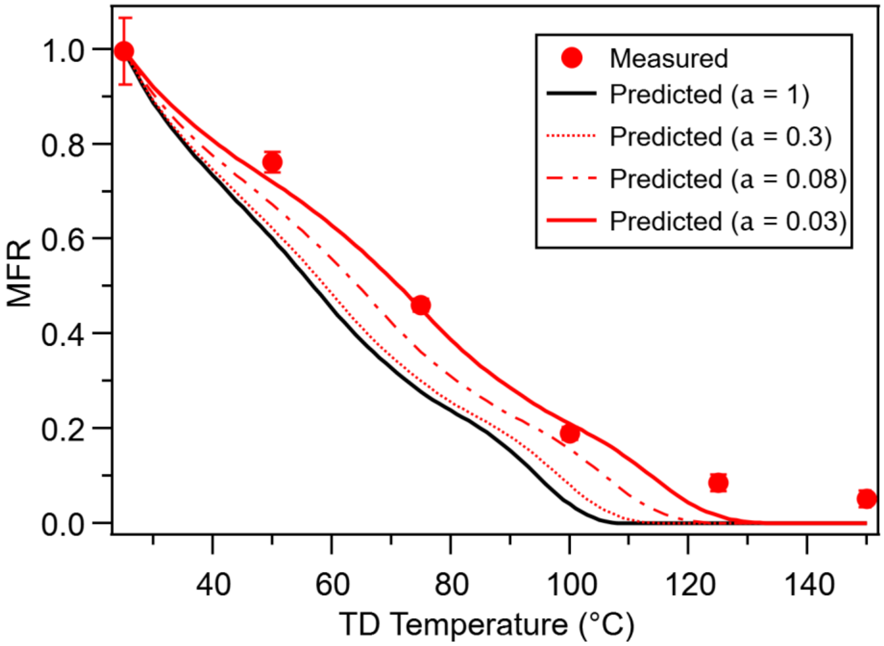
Q3. Figure 3: The author mentioned there is an increase in the average oxidation state (from -0.6 to -0.24) and a decrease in the average carbon number (from 5.2 to 4.9) with increasing photochemical aging (line 15 and 1ine 188-193 and Fig.3 caption). The decrease in carbon is not so significant, and from the figure, the mole fraction of carbon 7-9 compounds is even higher in the 1.7d photochemical aging condition than those in 10.7h condition. The conclusion of dominant fragmentation should be better elaborated. The SOA concentration actually increased with further oxidation (TableS1).

We agree that the decrease in average carbon # is not significant in this study. However, the effect of fragmentation is evident from an increasing fraction of smaller and more oxygenated compounds formed during photooxidation. For instance, the total fraction of C2-C7 SOA products increased from 81% in 10.7h SOA to 89% in 1.7d SOA. Of this fraction, the smaller carbon # compounds (C2-C4) which are indicative of fragmentation processes increased from 42% in 10.7h SOA to ~49% in 1.7d SOA. On the other hand, total fraction of >C7 (C8-C10) products declined from ~19% to ~11% as SOA aged. An increase in smaller and more oxygenated compounds, along with decrease in larger and less oxygenated products from 10.7h SOA to 1.7d SOA suggests that fragmentation reactions are responsible for the shift towards formation of smaller oxygenated compounds. We also use the term fragmentation to refer strictly to the decrease in carbon number, and not decrease in SOA concentration. Therefore, increase in SOA concentration is not strictly inconsistent with fragmentation, since the effect of increased $\overline{OSc}$ on volatility can exceed that of decreased carbon number. To clarify this, the following edits have been made to the manuscript in L193-198:

"…comprised of ~~long chain hydroxy acids~~ ~19% larger (C8–C10) and less oxygenated compounds, this fraction declined to ~11% at higher OH exposures. Furthermore, the total fraction of C2-C7 products increased from 81% to 89% when OH exposure increased from 10.7 h to 1.7 d. Of this fraction, the smaller carbon # compounds (C2-C4) which are indicative of fragmentation processes increased from 42% at 10.7 h to ~49% at 1.7 d. An increase in smaller and more oxygenated compounds, along with decrease in larger and less oxygenated products suggests that fragmentation reactions are responsible for the shift towards formation of smaller oxygenated compounds."

Q4. Figure 5 compares the measured and modeled mass thermograms for canola oil SOA. A mass accommodation coefficient of 0.03 was used in the model prediction. How sensitive is the predicted results to the accommodation coefficient? When using different accommodation coefficients of 1 and 0.03, the differences between MFR are only around 30% (at the same TD temperature). Maybe be better to add some thermograms using middle accommodation coefficient values between 1 and 0.03.

Thanks for the suggestion. We have modified Fig. 5 in the manuscript with different accommodation coefficients in the range 1-0.03.



Q5. Line 291-294: Other unidentified I/SVOCs may also play a role in the unexplained SOA mass.

Thanks for the suggestion. It is likely that other unidentified precursors such as IVOCs or SVOCs can contribute to the unexplained SOA. However, IVOCs from cooking emissions have not yet been positively identified. Based on the reviewer suggestion, the following changes have been made to the manuscript in L342-347:

"…model predictions. Furthermore, the unexplained SOA can likely arise from other unidentified S/IVOCs as hypothesized by Liu et al. (2017c). However, unlike traffic emissions (Zhao et al., 2014), S/IVOCs from cooking has not been positively identified. In addition, small VOC precursors like acrolein and malondialdehyde which have been measured in large quantities from cooking emissions (Klein et al., 2016a), may form SOA products having higher O/C ratios, which may better explain the O/C ratios observed in our experiments."

In addition, please also refer to response to Q7, Reviewer #2.

Q6. Line 314-321: When deriving the parameters for aldehyde oxidation, six tunable parameters were used to fit the measured SOA concentration. Compared with other systems, a lower mfrag was used in this study. The author attributed this to the greater fragmentation in this SOA system. However, the relative strength of these six parameters used to fit SOA concentrations is different for heptanal, 2-heptenal and 2,4- heptadienal oxidation experiments (Fig. S10). For example, mfrag used in heptanal experiment data is one order of magnitude lower than that used for the other two aldehydes. In addition to mfrag, what are the corresponding processes of the other five parameters? Are those related to gas-particle partitioning, functionalization, reactions with oxidants, or condensed-phase chemistry mentioned earlier (line 308-309)? It would be interesting to discuss the relationship between these parameters and their corresponding chemical or physical processes and how they behave in this system.

The six tunable parameters in SOM are: mfrag- which describes fragmentation reactions, ΔLVP- is decrease in logarithm of volatility upon addition of oxygen atom. Pox1, Pox2, Pox3, Pox4- describes addition of 1, 2, 3, and 4 oxygen atoms per reaction with OH, respectively.
The probability of a reaction with OH leading to fragmentation is calculated as $P_{frag} = (O:C)^{mfrag}$, where mfrag is the fitting parameter. The corresponding functionalization probability can be estimated as $P_{func} = 1 - P_{frag}$. Instantaneous gas/particle partitioning equilibrium is inherently assumed at every timestep in the model. Addition of oxygen atoms upon reaction with OH will govern the reaction with oxidants. Condensed-phase chemistry was not considered in the model.
To answer reviewer's question, as the # of oxygen atom addition increases on carbon backbone, it will likely result in the higher fragmentation probability. Therefore, as shown in Fig. S10, Pox4 is much higher for heptanal than 2-heptenal or 2,4-heptadienal, thereby having much lower mfrag value than 2-heptenal or 2,4-heptadienal.

Q7. In real cooking emissions, POA are also emitted with the aldehydes and other gasphase precursors. It would be useful if the authors can project how the inclusion of POA in the system would affect their results of O/C, etc.

Inclusion of POA during oxidation of cooking vapors will likely decrease the overall O:C (or $\overline{OSc}$) of the system as POA is less functionalized than SOA due to higher contributions from long chain fatty acids, such as C16, C18 thereby giving rise to $\overline{OSc}$ of POA from heated cooking oils ~ -1.7 (Takhar et al., 2019) which is much lower compared to $\overline{OSc}$ of cooking SOA measured in this study. Therefore, inclusion of POA would likely lead to an overall decrease in the average $\overline{OSc}$ or O:C of the system. However, it should be noted that POA can itself undergo heterogenous oxidation reactions in the atmosphere resulting in an increase in O:C. On the other hand, other gas phase precursors that can potentially contribute to total vapor emissions from cooking could be S/IVOCs which have not been positively identified from cooking emissions, but have been shown to contribute to SOA from other sources e.g. traffic (Zhao et al., 2014). A similar projection for cooking emissions can likely be made. Furthermore, depending on the cooking conditions, it has been shown that cooking can emit large amounts of terpenes upon addition of condiments or spices to heated cooking oils. Emissions of terpenes have been shown to significantly contribute to total SOA production (Klein et al., 2016b; Liu et al., 2017a). Based on the reviewer suggestion, the following edits have been made to the manuscript in L384-391:

"…Formation of SOA from cooking emissions in the atmosphere is likely influenced by emissions of POA, and other gas-phase precursors. Therefore, inclusion of POA during atmospheric processing of cooking emissions will likely influence the physicochemical properties of cooking SOA. For instance, with cooking POA being much less functionalized than SOA, inclusion of POA will likely decrease the system O:C (or $\overline{\mathrm{OSc}}$*). However, POA from cooking emissions can undergo heterogeneous reactions in the atmosphere, thereby increasing O:C (or* $\overline{\mathrm{OSc}}$*). On the other hand, there could potentially be contributions from other gas-phase precursors or S/IVOCs emitted from cooking vapors that can result in SOA formation. These precursors can potentially contribute to SOA formation from cooking emissions, but their oxidative evolution in the atmosphere is not well understood."*

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Anonymous Referee #2

The manuscript by Takhar et al. reported the SOA formation from OH oxidation of heated cooking oil and characterized the SOA composition and volatility using TD-GC/MS. A new method was used to analyze the composition of the complex product mixtures. The authors have found that an increase in OSc and a decrease in carbon number upon oxidation, which was attributed to fragmentation reactions during OH aging. In addition, by comparing product yields from individual precursors and applying an oxidation model, they conclude that aldehyde precursors are the main contributors of SOA formed from heated cooking oil. Overall, this is a well-written manuscript and the results can be a useful addition to better understanding the formation, composition and volatility of cooking SOA. I would recommend publication after a minor revision.

Specific Comments

1. The authors stated the increase in OSc and decrease in nC with oxidation. However, it is not very clear from Figure 3. Could the authors add the average OSc and nC in every panel of Figure 3?

Please refer to response to Q3, Reviewer #1.

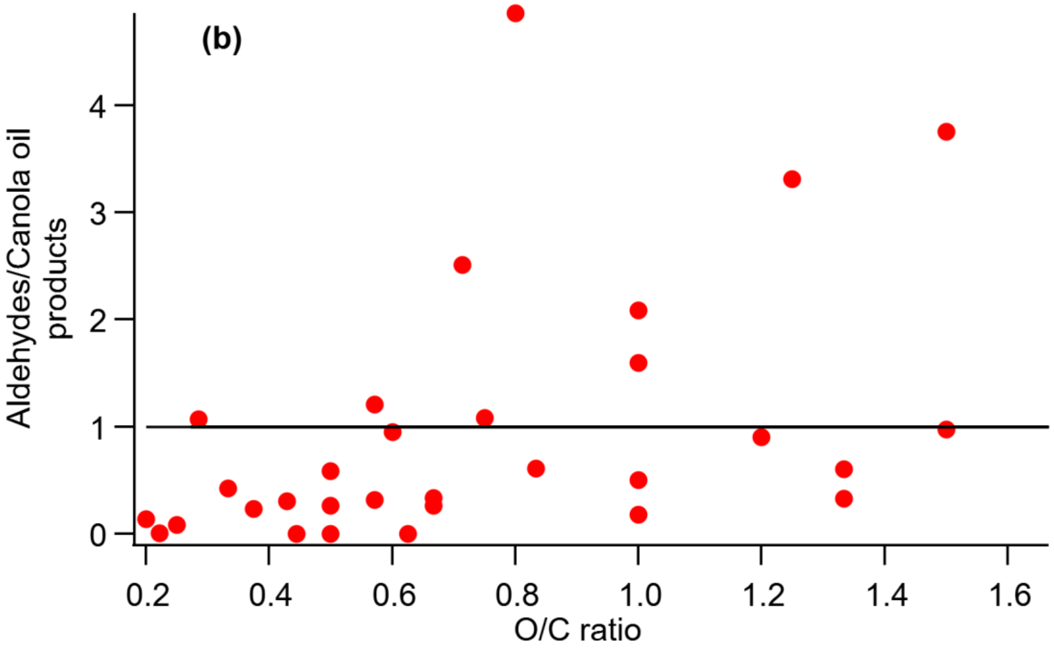
2. L193-197, the authors did not compare the same thing here. In Lambe et al. (2012, 2015), they found the evidence of fragmentation because of the decrease in SOA yield upon oxidation. However, in this study, the evidence of fragmentation is from the decrease in average carbon number. If the authors also look at the SOA yield, it never decrease with increasing photochemical age (Figure 7 and Table S1). Therefore, the authors could not state that fragmentation reactions happen earlier for cooking oil than other precursors (e.g., alkanes and isoprene).

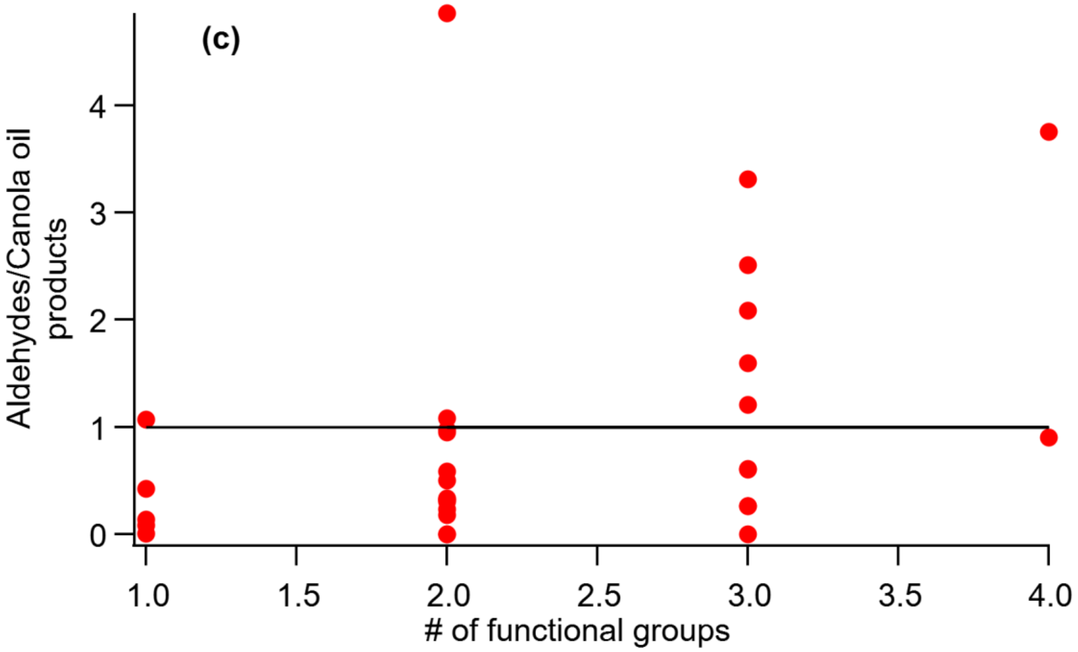
We agree that fragmentation reactions in this study are because of decrease in carbon # and not due to decrease in SOA yield. As suggested by the reviewer, L193-197 (now updated as L200-204) have been removed from the manuscript.

3. L293-295, I do not see any O/C or number of functional group dependence in Figure S9. The authors might need to change their way to present these data.

Figure S9(b) and (c) have been replaced with following figures, where y-axis is replotted as the ratio of aldehydes products to canola oil products instead of aldehydes products only. Since most of the products ratio lies below y-axis = 1 line suggesting that more oxygenated products partitions readily in canola oil SOA than individual aldehydes SOA. Please note that L293-295 is updated as L306-308.

Furthermore, Fig. S9(a) and (d) have also been updated with the similar y-axis against vapor pressure and carbon #.





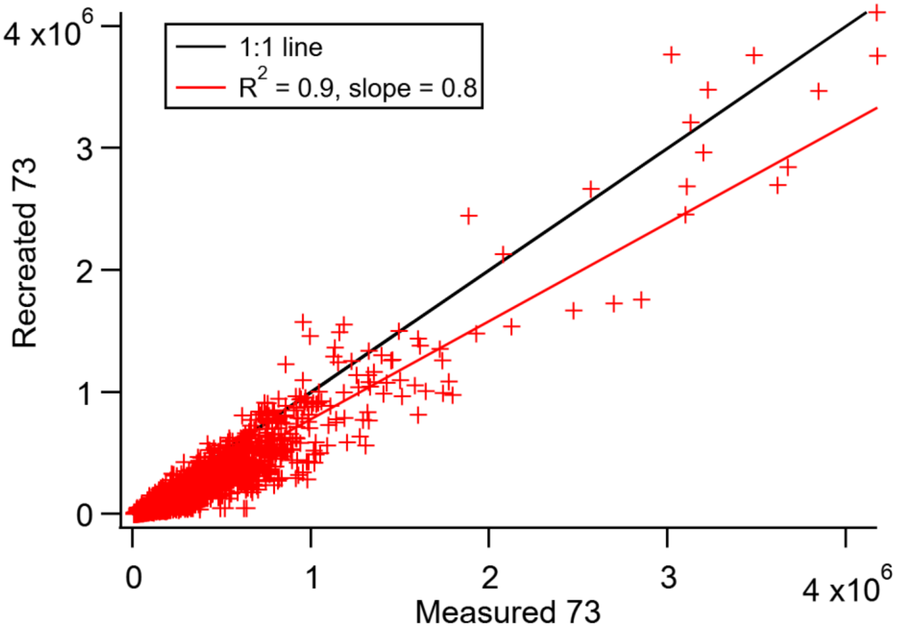
4. Figure 6 looks very interesting, but the authors did not discuss much about it. It seems that for higher-carbon-number products and 1-COOH products, the agreement is worse than others compounds. Any explanation for that?

The formation of higher carbon number products from the precursors photo-oxidized in this study were not observed likely due to negligible oligomerization reactions or reactions occurring in the particle-phase that can form SOA products with higher carbon # than the smaller carbon # parent VOCs photo-oxidized in this study. In addition, 1-COOH compounds are likely to be present as primary vapors in the gas-phase which can subsequently partition to the condensed phase upon SOA formation. Based on the reviewer comment, the following edits have been made to the manuscript in L302-305:

"…canola oil photooxidation. As shown in Fig. 6, the formation of higher carbon # products cannot be explained from the photooxidation of aldehydes used to predict oil oxidation products likely due to the assumption of negligible particle-phase or oligomerization reactions occurring in the condensed phase. In addition, higher carbon # acids are likely present as primary vapors in the gas phase which can then partition to the condensed phase upon SOA formation."

5. In Figure 2, the recreated m/z 73 seems agree well with the measured signal. Could the authors provide a scatter plot as well? Maybe it can replace Figure S4c.

Thanks for the suggestion. Fig. S4(c) is replaced with a scatter plot in the SI as shown below. However, it should be noted that a better assessment of recreated signal should be done by plotting chromatograms to evaluate any under/overestimation of the peaks or model compounds.



6. L201, what are the O/C ratios in previous studies? The author should put them here for comparison.

The O/C ratios reported in previous studies ranged between 0.24 to 0.46. The following edits have been made to the manuscript in L207-209:

"…The O:C ratios measured using an AMS (Kaltsonoudis et al., 2017; Liu et al., 2017b) ranged between 0.24-0.46 which are within a factor of 2 measured in this study."

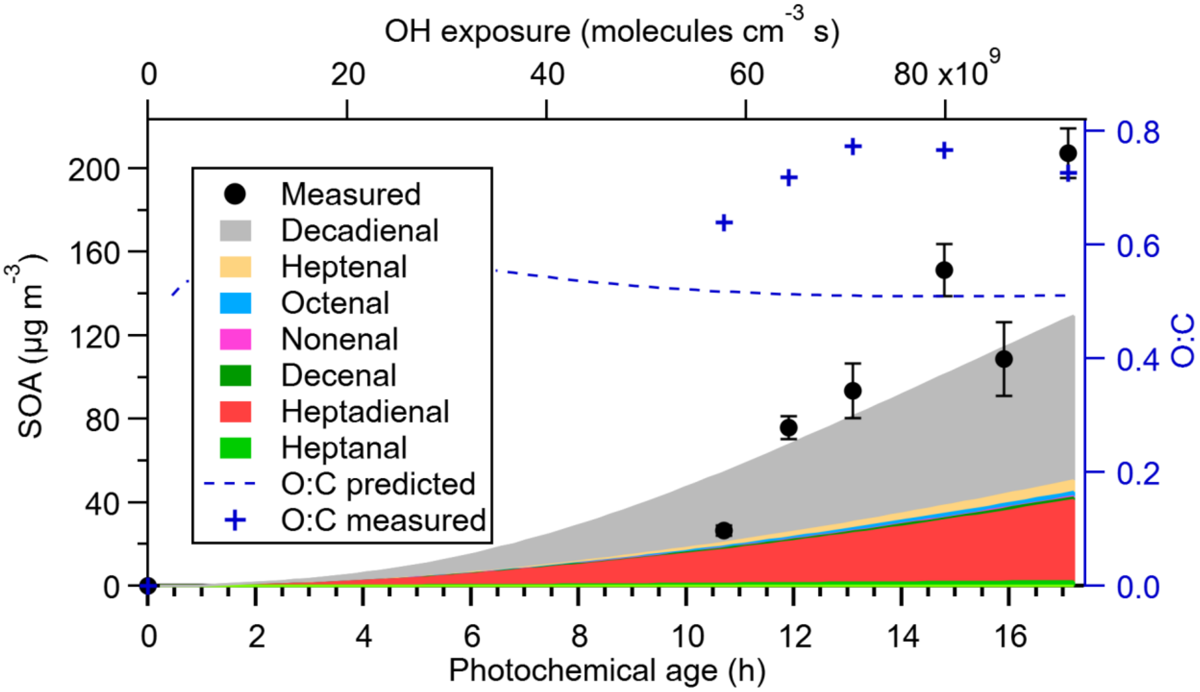
7. Exp. 7 (photochemical age = 40.7 h) is not included in Figure 7. The authors should mention that somewhere.

Thanks for pointing this out. The model simulations were run with OH exposure in the range similar to that of aldehyde photooxidation (Fig. S10). Upon further examination, we discovered that the OH exposure in the model results were incorrectly plotted in Fig. 7. In addition, we found that in earlier experiments we were unable to measure decadienal, but upon further examination, we found decadienal can be captured on Tenax tubes by measuring for longer duration, and so we have now included decadienal as one of the SOA precursors and estimated the formation of speciated products and total SOA using SOM. We have made corresponding changes to Figs. 6, 7 and S8. After correcting this error, the following corrections have been made to the manuscript with Fig. 7 revised as shown below. Our updated SOM estimate of heptadienal SOA contribution is now ~19% (down from 35% in our previous estimate, which corresponded to SOA at higher OH exposure). At the same time, decadienal accounts for 38% of the SOA. The total estimated contribution from aldehydes to canola oil SOA is 62%. Furthermore, it should be noted the modeled O:C estimate declined from 0.7 to 0.51 with the inclusion of decadienal.

In L334, "…aging conditions in the OH exposure range similar to that of aldehyde photooxidation."

In L338, "…Fig. 7, the model generally captures the amount of SOA formed to up to ~~within~~ ~~50~~62%, but…"

In L341-347 "…SOM predicts an O:C around 0.51~~7~~, which ~~lies~~ is within ~~±~~50~~20~~% of the measured O:C likely suggesting that the changes in chemical composition of cooking SOA is in ~~good~~ a reasonable agreement with the model predictions. Furthermore, the unexplained SOA can likely arise from other unidentified S/IVOCs as hypothesized by Liu et al. (2017c). However, unlike traffic emissions (Zhao et al., 2014), S/IVOCs from cooking has not been positively identified. In addition, small VOC precursors like acrolein and malondialdehyde which have been measured in large quantities from cooking emissions (Klein et al., 2016a), may form SOA products having higher O/C ratios, which may better explain the O/C ratios observed in our experiments."



Technical Corrections

L470-475, same references.

Thanks for pointing this out. Reference in line 473-475 (updated as L499-501) have been removed.

There is an incorrect number (1:2:1) in Figure S5.

Thanks for pointing this out. 1:2:1 has been corrected to 1.2:1.

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Characterization of secondary organic aerosol from heated-cooking oil emissions: evolution in composition and volatility

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Abstract. Cooking emissions account for a major fraction of urban organic aerosol. It is therefore important to understand the atmospheric evolution in the physical and chemical properties of organic compounds emitted from cooking activities. In this work, we investigate the formation of secondary organic aerosol (SOA) from oxidation of gas-phase organic compounds from heated cooking oil. The chemical composition of cooking SOA is analyzed using thermal desorption-gas chromatography-mass spectrometry (TD-GC/MS). While the particle-phase composition of SOA is a highly complex mixture, we adopt a new method to achieve molecular speciation of the SOA. All the GC elutable material is classified by the constituent functional groups, allowing us to provide a molecular description of its chemical evolution upon oxidative aging. Our results demonstrate an increase in average oxidation state (from -0.6 to -0.24), and decrease in average carbon number (from 5.2 to 4.9) with increasing photochemical aging of cooking oil, suggesting that fragmentation reactions are key processes in the oxidative aging of cooking emissions within 2 days equivalent of ambient oxidant exposure. Moreover, we estimate that aldehyde precursors from cooking emissions account for a majority of the SOA formation and oxidation products. Overall, our results provide insights into the atmospheric evolution of cooking SOA, a majority of which is derived from gas-phase oxidation of aldehydes.

1 Introduction

Organic aerosol (OA) has important impacts on air quality, climate and human health (Hallquist et al., 2009). OA is often composed of thousands of organic compounds formed from a variety of sources. In urban areas, particulate emissions from food cooking account for a significant fraction of OA (Allan et al., 2010; Crippa et al., 2013; Florou et al., 2017; Kostenidou et al., 2015; Lee et al., 2015; Mohr et al., 2012; Sun et al., 2011). Furthermore, volatile organic compounds (VOCs) are also emitted, and they can undergo oxidation and form secondary organic aerosol (SOA). Recent studies have reported the formation of SOA from meat charbroiling (Kaltsonoudis et al., 2017~~a~~) and heated cooking oils (Liu et al., 2017b, 2017c, 2018). Therefore, food cooking activities have substantial impacts on air quality in and downwind of urban areas.

The emission of VOCs from cooking is highly variable and depends on a number of factors such as cooking style, food, ingredients, and temperature (Fullana et al., 2004a, 2004b; Klein et al., 2016a, 2016b; Liu et al., 2017c; Schauer et al., 1999, 2002). Of the different classes of VOCs characterized in these studies, aldehydes have been shown to be the major group of VOCs emitted from cooking oils. These VOCs are chemically produced upon heating via peroxyl radical reactions of the fatty acids (Choe and Min, 2007; Gardner, 1989). Klein et al. (2016a) investigated the composition of nonmethane organic gas (NMOG) emissions from boiling, charbroiling, shallow and deep frying of various vegetables, meats, and cooking oils heated under different temperature conditions. The authors reported that emissions from shallow frying, deep frying and charbroiling are dominated by aldehydes, and the relative amounts depend on the type of oil used during cooking (Klein et al., 2016a). C7 aldehydes are the major species in emissions from canola oil, whereas C9 aldehydes are dominant from olive oil (Klein et al., 2016a). These differences in emission patterns of oils vary with composition of triglycerides present in the oil (Choe and Min, 2006). Katragadda et al. (2010) demonstrated up to an order of magnitude increase in emissions upon reaching the smoke point of cooking oils. In addition to emissions from cooking oil, the addition of condiments (herbs and peppers) to cooking leads to significant emissions of mono-, sesqui- and diterpenes in the gas phase (Klein et al., 2016b). Liu et al. (2017a) showed an order of magnitude increase in the emissions of VOCs when stir-frying with spices. Therefore, factors like cooking style, food, cooking temperature, and ingredients play a significant role in the chemical profile of cooking emissions (Fullana et al., 2004a, 2004b; Klein et al., 2016a, 2016b; Liu et al., 2017a, 2017c).

The VOCs emitted from cooking have been shown to produce significant amount of SOA rapidly in recent flow tube (Liu et al., 2017b) and smog chamber studies (Kaltsonoudis et al., 2017~~a~~; Liu et al., 2017c, 2018). Kaltsonoudis et al. (2017~~a~~) and Liu et al. (2017b, 2018) showed an increase in O/C ratio upon a few hours of atmospheric aging suggesting lightly oxidized cooking SOA. Furthermore, Liu et al. (2017b) showed significant production of SOA with increasing OH exposure for different cooking oils. Thus far studies have only focused on formation potential of SOA from cooking emissions. Despite high emission rates of VOCs from cooking, the understanding of SOA composition from these emissions remains limited.

Source apportionment using aerosol mass spectrometry (AMS) data in urban areas has often revealed a Cooking Organic Aerosol (COA) factor, but it is unclear how this factor is related to cooking emissions. Many studies reported that the mass spectra associated with this factor resemble that of hydrocarbon-like organic aerosol (HOA) factor from other non-cooking sources (Dall'Osto et al., 2015; Hayes et al., 2013; Huang et al., 2010; Mohr et al., 2009, 2012). In addition, it is often unclear whether ambient COA represents primary or secondary organic aerosol from cooking emissions (Dall'Osto et al., 2015; Florou et al., 2017; Kaltsonoudis et al., 2017~~b~~; Kostenidou et al., 2015). Laboratory studies (Liu et al., 2017b, 2018) showed that the mass spectra for primary cooking organic aerosol exhibited strong correlation with ambient COA factor (Lee et al., 2015), but the cooking SOA mass spectra showed some similarities to ambient semi-volatile oxygenated OA (SV-OOA) factor. These measurements highlight the challenges in assigning COA factor without understanding the changes in chemical composition occurring during oxidation of cooking emissions.

In general, there is a need to better understand the molecular composition contributing to aged COA. In this study, we investigate detailed chemical composition of cooking SOA at the molecular level. The objectives of this study are to: (i) understand the detailed chemical speciation of cooking SOA using TD-GC/MS, (ii) describe chemical evolution in SOA upon atmospheric aging, and (iii) attribute formation of SOA to different VOCs emitted from food cooking emissions. In this work we use heated cooking oil as a model for food cooking emissions. We show that the majority of the SOA is derived from oxidation of aldehydes, and the oxidation mechanisms are dominated by fragmentation reactions. Overall, our results provide useful insights into the evolution of cooking SOA, which may be incorporated into chemical transport models for better predicting OA formation from cooking emissions in the atmosphere.

2 Experimental methods

2.1 Flow tube experiments

The experimental setup is shown in Fig. 1, and experimental conditions are listed in Table S1. For each experiment, 30-40 mL of canola oil was heated at 250 °C on an electric heating plate in a Pyrex bottle resulting in an average cooking oil temperature of 180 °C, as measured by a thermocouple in direct contact with the heated oil. Purified air flowed over the headspace of the heated oil at a rate of 0.2 L min^{-1} and then diluted by a factor of 50. 0.2 L min^{-1} of the total diluted flow was passed through a Teflon filter to remove particles, and the oil vapors were introduced into a custom-built 10 L quartz flow tube reactor. A separate flow of oxygen (99.6%) was irradiated in a UV ozone generator (UVP 97006601) to produce ozone and was also introduced into the flow tube reactor. In parallel, purified air was flowed through a water bubbler into the reactor to provide water vapor. The combined flow rate through the flow tube was set at 3 L min^{-1}, resulting in an average residence time of approximately 200 s.

In the flow tube, hydroxyl radicals were produced through the photolysis of ozone irradiated by a UV lamp (λ = 254 nm) in the presence of water vapor. The integrated OH exposure was measured indirectly from the loss of cyclopentane which was monitored by a gas chromatography-flame ionization detector (GC-FID, model 8610C, SRI Instruments Inc.) equipped with a Tenax TA trap sampling downstream of the flow tube at a rate of 0.15 L min^{-1}. In this study, the experiments were conducted at different OH exposures ranging from 5.77×10^{10} to 2.2×10^{11} molecules cm^{-3} s. OH exposure in this range is equivalent to ~11 to 41 h of atmospheric oxidation, respectively, assuming a 24-h average atmospheric OH concentration of 1.5×10^{6} molecules cm^{-3} (Mao et al., 2009). The effect of ozone on the SOA formation was found to be negligible as the reaction timescales of aldehydes with ozone were calculated to be at least 100 times longer than those with OH. A sample calculation for methacrolein reaction timescales with OH and ozone is shown in SI in Sect. 1.

Downstream of the flow tube, pre-baked quartz fiber filter and Tenax tube samples were collected for offline chemical analysis. The changes in the particle size distribution and volume concentration were monitored using a scanning mobility particle sizer (SMPS) with a differential mobility analyzer (TSI 3081), and a condensation particle counter (TSI 3781). A constant density of 1.4 g cm^{-3} was assumed to convert particle volume concentration into mass concentration (Chan et al., 2010). Relative humidity and temperature were monitored by an Omega HX94C RH/T transmitter and were maintained at 65-70%, and 19-20 °C, respectively for all experiments. A fast stepping/scanning thermodenuder (TD, Aerodyne Inc. Billerica, USA) was also placed downstream of the flow tube to measure SOA evaporation rates. Details about TD operating conditions and analysis can be found in Takhar et al. (2019). The TD was only operated during one experiment in which the OH exposure was 9.23×10^{10} molecules cm^{-3} s. The SOA was systematically heated in a TD from 25 °C to 175 °C, and changes in particle volume concentrations and corresponding mass fraction remaining (MFR) were measured using a SMPS. The SOA size distribution during TD operation and volatility distribution are shown in Fig. S1 and S2, respectively. A kinetic mass transfer model developed by Riipinen et al. (2010) was used to interpret the TD data. The inputs to the model are volatility distribution of OA, enthalpy of vaporization, and mass accommodation coefficients. Compound groups are translated into volatility distributions by binning components according to their saturation concentrations (Donahue et al., 2006). Parameterization for enthalpy of vaporization was similar to that of Takhar et al. (2019). We assume a surface tension of 0.05 N m^{-1}, gas-phase diffusion coefficients of 5×10^{-6} m^2 s^{-1} for all simulations similar to that reported in Riipinen et al. (2010).

2.2 Chemical characterization of SOA

Tenax tube and quartz filter samples were analyzed separately by thermal desorption gas chromatography mass spectrometry (TD-GC/MS) for detailed chemical speciation of gas- and particle-phase organic compounds. The analyses were performed using a thermal desorption system (TDS 3, Gerstel) combined with a gas chromatography (7890B, Agilent)-mass spectrometer (5977A, Agilent). For gas-phase analysis, concentrations of aldehydes (C7 to C10 *n*-alkanals, alkenals and alkadienals) collected on Tenax tube samples before photooxidation (downstream of the flow tube, with lights off) were quantified. For particle-phase analysis, thermal desorption of quartz filters was performed with *in situ* derivatization using *N*-trimethylsilyl-*N*-methyl trifluoroacetamide (MSTFA). A known amount of deuterated 3-hydroxy-1,5-pentanedioic-2,2,3,4,4-d_5 acid, and *n*-pentadecane-d_{32} (CDN isotopes) was injected, respectively, onto quartz filter punches, and Tenax tubes as internal standards before the samples were desorbed in the TDS. All GC/MS analysis was performed using a non-polar DB5 column (Rxi-5Sil MS, Restek). Details of the operating parameters (GC column, GC and TDS temperature ramps, MS parameters) can be found in Sect. ~~1~~ 2 of SI.

With *in situ* derivatization, polar organic compounds react rapidly with MSTFA at elevated temperatures during thermal desorption, and functional groups with acidic hydrogen atoms (such as –OH) are replaced by a less polar trimethylsilyl (TMS, [$-OSi(CH_3)_3$]) group. This reduction in polarity allows the derivatized analyte to elute from a non-polar column and analyzed by subsequent electron impact (EI) at 70 eV. Derivatized compounds produce a signature fragment ion at mass-to-charge (*m/z*) 73 ($-Si(CH_3)_3^+$) arising from the scission of O-Si bond in $R-O-[Si(CH_3)_3]$. In other words, all derivatized compounds produce ions with *m/z* 73 during analysis. Therefore, the total signal at *m/z* 73 can be taken as the total concentration of organic compounds with at least one hydroxyl group (including both –OH and –C(O)OH) present in cooking SOA, much like how *m/z* 57 represents total concentration of aliphatic compounds in hydrocarbon mixtures (Zhao et al., 2014, 2015). It should be noted that organic peroxides (R–OOH) were also found to be derivatized, but the major reaction product formed is $R-O-[Si(CH_3)_3]$ (which is also formed from R–OH derivatization) as shown in Fig. S3. Here we assume alcohols and acids are the major components, but will explore the potential role of ROOH on the overall chemical composition in Sect. 3.1.

As shown in Fig. 2, many compounds in cooking SOA contain at least one –OH group and the chromatogram of *m/z* 73 is typical of that for a highly complex mixture or unresolved complex mixture (UCM). Using traditional analytical techniques like GC/MS it is difficult to deconvolute the UCM. However, knowledge about mass spectral fragmentation of TMS derivatives can be used to understand the compounds contributing to the UCM. Table S2 shows a list of compounds containing multiple functional groups e.g. -COOH, -OH resulting in different combinations of compound classes like dicarboxylic acids, hydroxy acids, hydroxy dicarboxylic acids, and dihydroxy dicarboxylic acids with different carbon numbers. As mentioned earlier, we acknowledge the potential contribution from ROOH, but will first assume the functional groups shown in Table S2 here, and consider ROOH in more detail in a later section. The compound groups shown in Table S2 are expected to be formed from oxidation of aldehydes and be derivatized by MSTFA. The TMS derivatives of these compounds share common ion fragments in their EI mass spectra: *m/z* 73 $[Si(CH_3)_3]^{\cdot +}$, 75, 147 $[(CH_3)_2Si{=}O(CH_3)_3]^{\cdot +}$, M-15 $[M-CH_3]^{\cdot +}$ (Jaoui et al., 2004, 2005; Yu et al., 1998). Most importantly, all TMS derivatives exhibit quantifiable peaks at *m/z* 73 (ubiquitous ion for all derivatives) and M-15 (ion specific to each compound group, hereby referred to as the pseudo-parent ion). We also obtained the characteristic ratio of these two ions for each compound group ($f_{M-15/73}$) from NIST mass spectral libraries and from analyzing authentic standards. To verify the validity of this method, we calculate the total *m/z* 73 ion signal that is attributable to these compound groups by taking the chromatograms of the pseudo-parent ion for each compound group, dividing by its characteristic ratio $f_{M-15/73}$ and then summing across all compound groups as shown in Eq. (1).

$$S_{73,t}^{sum} = \sum_i \frac{S_{M-15,i,t}}{f_{M-15/73,i}} \tag{1}$$

where $S_{73,t}^{sum}$ is the *m/z* 73 ion signal at retention time *t* that is attributable to all compound groups listed in Table S2, $S_{M-15,i,t}$ is the signal of the pseudo-parent ion for compound group *i* at retention time *t*, $f_{M-15/73,i}$ is the characteristic ratio of pseudo-parent ion to *m/z* 73. This approach is similar to that described in Isaacman-VanWertz et al. (2020). As shown in Fig. 2, $S_{73,t}^{sum}$ shows excellent agreement with the measured *m/z* 73 ion signal, suggesting that the *m/z* 73 signal, which is representative of all TMS derivatives, is almost entirely comprised of contributions from the compound groups listed in Table S2. This agreement between our bottom-up approach and measured signal provides confidence that our method is able to provide information about the chemical composition of highly complex mixture.

With the signals from all the pseudo parent ions for all compound groups, the total mass of each compound group was then calculated using Eq. (2).

$$M_i = \frac{TA_i}{RF_i} \text{x} \frac{1}{f_{M-15/73,i}} \tag{2}$$

where, M_i is the mass of compound group i, TA_i is the total integrated signal of pseudo-parent ion for compound group *i* (normalized by the signal of deuterated internal standard), RF is the response factor (calculated from calibration curves of fatty acids and dicarboxylic acids authentic standards) of compound group *i*, and $f_{M-15/73,i}$ is the characteristic ratio of pseudo-parent ion to *m/z* 73 for compound group *i*. A more detailed, step-by-step description of the procedure can be found in the SI in Sect. ~~2~~3, and illustrated in Fig. S4 with corresponding uncertainties in the fitting procedure shown in Fig. S5.

3 Results and discussion

3.1 Chemical evolution of SOA

As described in Sect. 2.2, components in cooking SOA were classified by functional groups and carbon number. To describe the overall changes in SOA composition with increasing OH exposure, we use the average carbon oxidation state ($\overline{\mathrm{OSc}}$) as a metric for the evolving composition of a complex mixture undergoing oxidation (Kroll et al., 2011). Both $\overline{\mathrm{OSc}}$ and number of carbon atoms (nc) for each compound group are calculated from the GC-derived chemical composition. The total mole fraction of C, H and O was calculated for each sample which was then used to calculate the bulk $\overline{\mathrm{OSc}}$ using the Eq. $2 \times O:C - H:C$ (Kroll et al., 2011). The evolution in this framework for canola oil SOA is shown in Fig. 3. The bulk $\overline{\mathrm{OSc}}$ was observed to increase from -0.6 to -0.24 when OH exposure increased from 5.77 to 22.0×10^{10} molecules cm^{-3} s for canola oil SOA. For comparison, Liu et al. (2017b) showed an initial decrease in $\overline{\mathrm{OSc}}$ and O:C, but gradually stabilized at OH exposure greater than 9×10^{10} molecules cm^{-3} s. For the $\overline{\mathrm{OSc}}$ range reported here, the $\overline{\mathrm{OSc}}$ of cooking SOA falls in the range of SV-OOA as determined from factor analysis of AMS data (Canagaratna et al., 2015). This degree of oxygenation is greater than that of the COA factor measured by AMS, which is reported to be around -1.37 (Canagaratna et al., 2015). This difference suggests that the COA factor resolved using PMF analysis is likely of primary origin and does not represent SOA formed from atmospheric oxidation of cooking emissions. Furthermore, previous GC/MS analysis showed for POA from cooking oils, an $\overline{\mathrm{OSc}}$ of -1.66 (canola oil) and -1.7 (beef tallow, olive oil) was calculated (Takhar et al., 2019). These observations again suggest that COA factor measured by AMS ~~is derived of~~represents primary cooking emissions.

In addition to carbon oxidation state, knowledge about molecular composition provides further insights into the oxidation mechanisms. Canola oil SOA at an OH exposure of 5.77×10^{10} molecules cm^{-3} s is comprised of ~~long chain hydroxy acids~~ ~19% larger (C8-C10) and less oxygenated compounds, this fraction declined to ~11% at higher OH exposures. Furthermore, the total fraction of C2-C7 products increased from 81% to 89% when OH exposure increased from 10.7 h to 1.7 d. Of this fraction, the smaller carbon # compounds (C2-C4) which are indicative of fragmentation processes increased from 42% at 10.7 h to ~49% at 1.7 d. An increase in smaller and more oxygenated compounds, along with decrease in larger and less oxygenated products suggests that fragmentation reactions are responsible for the shift towards formation of smaller oxygenated compounds. As a result, oxidation simultaneously leads to higher $\overline{\mathrm{OSc}}$ and lower carbon number on average. Based on the compounds observable by our technique, this trend suggests that fragmentation reactions are key processes in the oxidative evolution of cooking emissions. ~~These findings suggest an early onset of fragmentation reactions upon atmospheric aging of cooking emissions contrary to other SOA systems, such as alkanes and isoprene (Lambe et al., 2012, 2015), in which fragmentation reactions dominate at later OH~~

~~exposures (>5×10^{11} molecules cm^{-3} s). Therefore, predicting OA concentrations from cooking emissions would require earlier fragmentation of SOA in climate and air quality models.~~

The compounds observed here can also be compared to previously measured bulk composition using elemental ratios, such as those presented in a Van Krevelen (VK) diagram (Heald et al., 2010). As shown in Fig. 4, the O:C ratio in our study ranged between 0.64 and 0.79 when OH exposure increased from 5.77×10^{10} to 22.0×10^{10} molecules cm^{-3} s. The O:C ratios measured using an AMS (Kaltsonoudis et al., 2017~~a~~; Liu et al., 2017b) ranged between 0.24-0.46 which are within a factor of 2 measured in this study. ~~These ratios are within a factor of 2 than previously reported AMS measurements of cooking oil SOA (Kaltsonoudis et al., 2017a; Liu et al., 2017b).~~ Furthermore, the H:C versus O:C trend is linear with a slope of -0.19, which lies between the slope of 0 measured for low-NO$_x$ oxidation reported by Liu et al. (2017b) and -0.4 for high-NO$_x$ conditions (Liu et al., 2018). Therefore, based on elemental ratios, the evolution in SOA composition measured in this study is comparable to that in bulk average properties estimated by AMS. Furthermore, we use 2D-VBS framework developed by Donahue et al. (2012) to investigate OA chemistry, and understand the evolution of cooking SOA through changes in the volatility of SOA system. The vapor pressures of the identified compounds are calculated using group contribution method (Pankow and Asher, 2008) where experimentally determined vapor pressures were unavailable, and reported in Table S2. The observed compounds in SOA have a broad range of volatilities, since they were formed from oxidation of a complex ensemble of VOC precursors. As shown in Fig. S6, there is minor decrease in overall volatility of the mixture (change lies within one decade in C*) irrespective of the presence of peroxides, while $\overline{\text{OSc}}$ is increasing with oxidation. This increase in oxidation state is coincident with increasing fragmentation upon oxidation, and, as a result, the overall change in the bulk volatility of canola oil SOA is relatively small.

As mentioned earlier in Sect. 2.2, there is a potential to misclassify ROOH as ROH using our current GC/MS method. In Fig. S3, we show that derivatization of cumene hydroperoxide forms the TMS of hydroxy-cumene in our system. Here we further examine the chemical composition by assuming that each -O-[Si(CH$_3$)$_3$] group observed originates from an -OOH group in the SOA.~~, and to support this argument we show that derivatization of cumene hydroperoxide (Sigma Aldrich Co.) is observed as TMS of hydroxy-cumene in our system as shown in Fig. S3.~~ It should be noted that replacing –OH with –OOH results in a higher estimate of O:C (and $\overline{\text{OSc}}$) but does not change H:C or carbon #. Furthermore, since pseudo molecular ion fraction ($f_{M-15/73}$) for organic peroxides (needed for quantification) is unknown, we assume that it is similar to those presented in Table S2. As shown in Fig. S7, if all observed –OH groups are –OOH groups, the VK-slope ~~is~~would be -0.15, which is similar to -0.19 calculated based on the no-peroxide assumption. Similarly, Fig. S6 shows that this uncertainty in hydroxyl group identification has negligible effect on estimation of vapor pressure or volatility in the 2D-VBS framework. Therefore, this potential misclassification of peroxide groups may lead to an underestimation in O:C and $\overline{\text{OSc}}$, but is not expected to affect estimates of volatility and our general conclusions about the importance of fragmentation reactions. In the future, analytical techniques such as extractive electrospray ionization time-of-flight mass spectrometry (Lopez-Hilfiker et al., 2019) may be useful to better understand the composition of peroxides from cooking SOA. While the misclassification of peroxides may have little impact on the bulk properties such as average O:C ratios, there may be important implications on understanding the reactivity of the SOA.

3.2 Evaporation rates of SOA

The volatility of the SOA is also probed by measuring the evaporation rates in a heated thermodenuder and compared to the rates expected from the measured composition. In order to derive the evaporation rates from the measured chemical composition of cooking SOA, we use the kinetic mass transfer model developed by Riipinen et al. (2010). Among the inputs into the model, the mass accommodation coefficient is a critical but uncertain parameter that accounts for the mass transfer limitations in the system.

Figure 5 shows both measured and modeled mass thermograms for canola oil SOA. We observe that for canola oil SOA, mass accommodation coefficient of 0.03 is needed to predict the experimentally determined mass thermograms. An accommodation coefficient of <1 suggests that mass transfer limitations in the system likely occurring in the condensed~~-~~ phase. Formation of multifunctional organic compounds such as those observed in this study is likely responsible for an increase in viscosity through increasing hydrogen bonding and other polar interactions (Rothfuss and Petters, 2016). It should be noted that Takhar et al. (2019) reported similar magnitudes of mass accommodation coefficients for heterogeneous oxidation of cooking oil particles. Due to similarity in the type of functional groups present in both aging pathways, we believe the decrease in mass accommodation coefficients for both systems undergo similar changes in phase and/or viscosity.

These measurements of evaporation rates are consistent with the volatilities expected from our measured composition of SOA containing small oxygenated compounds. Although mass accommodation coefficients are highly uncertain, the mass accommodation coefficients for other SOA systems have been measured to be even lower on the order of 10^{-4} (Cappa and Wilson, 2011), which would require the volatilities to be even higher to explain the measured evaporation rates. Therefore, the TD measurements support the conclusion that smaller oxygenated compounds are produced from oxidation of cooking oil vapors, and that fragmentation reactions are dominant. Furthermore, these measurements provide useful inputs into chemical transport models for predicting SOA formation and gas-particle partitioning. Our previous work (Takhar et al., 2019) showed that even at $\alpha = 10^{-2}$, gas-particle partitioning timescales are short (within hours) and the assumption of equilibrium partitioning still holds for regional scale SOA formation. Further work is needed to directly measure the viscosity of cooking SOA, and corresponding mixing timescales to better constrain the physicochemical properties of cooking SOA.

3.3 Contribution of aldehydes to observed oxidation products and total SOA

Since cooking oil vapors are comprised of a number of reactive aldehydes that can lead to SOA formation, we conduct further experiments of SOA formation from these precursors and identify the relative contributions to observed oxidation products and to total SOA. These results are applied to the heated cooking oil experiments to understand the role of aldehydes in the overall production and evolution of cooking oil SOA.

3.3.1 Formation of particle-phase oxidation products

As described in the earlier sections, we are able to quantify the mass concentrations of different compound groups (6 different combinations of functional groups, from C2 to C~~9~~10, summarized in Table S2) in the particle phase for all experiments. We denote the observed mass concentrations of compound group *i* in SOA from canola oil photooxidation as M_i^{oil}. The expected precursors to these oxidation products are likely aldehydes, since aldehydes are emitted in significant amounts and are highly reactive. To examine this hypothesis, here we calculate the formation of these observed compound groups from oxidation of aldehydes. For this calculation, heptanal, *trans*-2-heptenal, *trans*-2-octenal, ~~and~~ *trans*,*trans*-2,4-heptadienal, and *trans*,*trans*-2,4-decadienal (Sigma Aldrich Co.) were considered because these aldehydes are the dominant VOC precursors emitted from heated canola oil in our experiments as shown in Fig. S8. ~~Unlike previous work by Fullana et al. (2004b) and Klein et al. (2016a), gas-phase concentrations of decadienals were minimal in our experiments.~~ More volatile aldehydes, such as acrolein and methacrolein, were likely present but could not be captured and analyzed by our techniques. The molar amount reacted for each aldehyde *j* in the canola oil oxidation experiments is denoted as ΔVOC_j^{oil}, and was calculated based on the measured OH exposure.

In order to estimate the contribution from oxidation of an aldehyde *j* in the gas-phase mix to the formation of each compound group *i*, we conducted a series of experiments in which a representative aldehyde was oxidized, and the molar yields of the various compounds were measured:

$$\gamma_{ij} = \frac{M_{ij}^{ind}/MW_i}{\Delta VOC_j^{ind}} \tag{3}$$

where γ_{ij} represents the molar yield of compound group *i* from precursor *j*, M_{ij}^{ind} denotes the mass concentration of compound *i* observed in photooxidation experiments in which aldehyde *j* was the sole precursor, MW_i is the molecular weight of compound *i*, and ΔVOC_j^{ind} is the amount of precursor *j* reacted in each experiment. γ_{ij} is then applied to the heated cooking oil experiments to estimate the amount of oxidation products that would form from each precursor:

$$M_i^{sum} = \sum_j \gamma_{ij} \Delta VOC_j^{oil} MW_i \tag{4}$$

A sample calculation for this analysis is presented in Sect. ~~3~~ 4 of SI. The comparison between M_i^{sum} (contribution of aldehyde oxidation to formation of compound *i*) and M_i^{oil} (observed concentrations of compound *i*) is shown in Fig. 6. Based on this methodology, oxidation of aldehydes accounts for ~~5~~63 µg m^{-3} (M_i^{sum}) of the observed 75 µg m^{-3} (M_i^{oil}) (or 84~~75~~%) particle-phase oxidation products measured at an OH exposure of 6.43×10^{10} molecules cm^{-3} s. The contributions of alkanals (heptanal), alkenals (heptenal + octenal) and alkadienals (heptadienal + decadienal) are 7%, ~31% and 46~~37~~%, respectively.

While the amount of oxidation products expected from aldehydes is somewhat lower than that observed in canola oil SOA, this difference may arise from differences in gas-particle partitioning between single aldehyde photooxidation and canola oil photooxidation. As shown in Fig. 6, the formation of higher carbon # products cannot be explained from the photooxidation of aldehydes used to predict oil oxidation products likely due to the assumption of negligible particle-phase or oligomerization reactions occurring in the condensed phase. In addition, higher carbon # acids are likely present as primary vapors in the gas phase which can then partition to the condensed phase upon SOA formation. As shown in Fig. S9, more oxygenated compounds (higher O:C and greater number of functional groups) tend to be more abundant in the canola oil SOA than expected from aldehyde photooxidation, suggesting that canola oil SOA is more favorable for oxygenated compounds to partition than SOA from individual aldehydes. On the other hand, there is no clear trend in partitioning with respect to vapor pressures and carbon number. It should be noted that uncertainties in the fitting procedure or estimation in the pseudo molecular ion (refer to Table S2 and Fig. S5) can also result in uncertainties between -40% and +20%. Therefore, in summary, the quantified oxidation products from canola oil SOA are generally consistent with those from aldehyde photooxidation, and the relative amounts may be subject to further changes due to gas-particle partitioning.

3.3.2 Using the statistical oxidation model (SOM) framework

To further explore the evolution of canola oil SOA, we applied our results to the statistical oxidation model (SOM) framework developed by Cappa and Wilson (Cappa et al., 2013; Cappa and Wilson, 2012). SOM describes the oxidation chemistry of a VOC precursor through multi-generational space defined by the number of carbon and oxygen atoms present in the precursor and its possible SOA product molecules. The SOM does not specifically track the product composition in terms of functional groups, but provides adequate details to represent key atmospheric processes such as gas-particle partitioning, fragmentation, functionalization, reactions with oxidants, condensed-phase chemistry. The model has been applied to chamber experiments to derive parametrizations by fitting experimental data to both SOA mass concentration and the bulk aerosol O/C ratio. Eluri et al. (2018) used the chamber derived parameterizations to predict the properties of SOA generated from diesel exhaust in an oxidation flow tube reactor.

To the best of the authors' knowledge, there are no parameterizations for the oxidation of aldehydes. Therefore, in this study we first derived the parameterizations for aldehyde oxidation, and then use these parameters to predict the SOA mass concentrations. In order to obtain the parameters, we fit the measured SOA concentration from oxidation of heptanal, *trans*-2-heptenal, *trans,trans*-2,4-heptadienal at different OH exposures to optimize the six tunable parameters under low-NO_x conditions (shown in Fig. S10). Best fit SOM parameters indicate that photooxidation leads to fragmentation per reaction with OH, as shown by a lower *mfrag* than compared to other systems e.g. alkanes (≥2 for branched, cyclic or *n*-alkane under low-NO_x conditions (Eluri et al., 2018)). Since a lower value for *mfrag* represents greater fragmentation (Cappa and Wilson, 2012), this again reflects the higher propensity for fragmentation in this SOA system. The optimized parameters were then used to predict the SOA concentration for canola oil photooxidation under different aging conditions in the OH exposure range similar to that of aldehyde photooxidation.

Based on these established parameterizations for different aldehydes, model simulations were conducted for canola oil having a mixture of aldehydes under different photochemical aging conditions. It should be noted that we used parameterizations of heptanal for all alkanals, heptenal for all alkenals, and heptadienal for alkadienals. As shown in Fig. 7, the model generally captures the amount of SOA formed to up to ~~within~~ ~~50~~62%, but overpredicts SOA formation at lower photochemical ages and underpredicts SOA concentrations at higher photochemical ages. In addition, SOM also tracks atomic O/C ratio which were further compared with the measured O/C ratio. SOM predicts an O:C around 0.51~~7~~, which ~~lies~~ is within ~~±~~50~~20~~% of the measured O:C likely suggesting that the changes in chemical composition of cooking SOA is in ~~good~~a reasonable agreement with the model predictions. Furthermore, the unexplained SOA can likely arise from other unidentified S/IVOCs as hypothesized by Liu et al. (2017c). However, unlike traffic emissions (Zhao et al., 2014), S/IVOCs from cooking has not been positively identified. In addition, small VOC precursors like acrolein and malondialdehyde which have been measured in large quantities from cooking emissions (Klein et al., 2016a), may form SOA products having higher O/C ratios, which may better explain the O/C ratios observed in our experiments.

One inconsistency between the model and measurements is the slope at which SOA is being formed. The experimental data suggest a steeper trend of SOA formation while the model predicts a more gradual increase in SOA formation. A potential explanation for this discrepancy is the contribution from other unmeasured VOCs. These VOCs are less reactive than those considered in the model, such that they contribute to higher SOA at higher OH exposures. Alternatively, these missing VOCs are more volatile such that more of their SOA is formed at later generations of oxidation. For example, acrolein forms SOA with measurable yields (Chan et al., 2010) and is emitted at large amounts from heated cooking oils (Klein et al., 2016a). Despite these limitations, these parameterizations generally capture the amount of SOA formed and its degree of oxidation (O/C) on oxidation timescales relevant to urban areas (within 2 days) and are useful for representing cooking oil emissions in the chemical transport models. Overall, the amount of SOA formed and the evolution upon oxidation can be well described by photooxidation of aldehydes.

4 Conclusions and implications

In this work, we characterized the detailed chemical composition of SOA generated from cooking oil vapors. We showed that cooking SOA occurring as highly complex mixture can be deconvoluted using mass spectral fragmentation pattern to extract useful information about the chemical identities of organic compounds, such as functional groups and carbon number. Using this detailed chemical composition of cooking SOA, we showed that fragmentation is an important pathway for oxidative processing of cooking emissions in the atmosphere even within short timescales of oxidation. Furthermore, we showed that aldehydes can reasonably explain the formation of SOA generated from cooking oil vapors and the oxidative evolution as described using a multi-generational oxidation model. Our study, therefore, highlights the importance of molecular composition in constraining the chemical properties of cooking SOA, as well as understanding the contribution of aldehydes in formation of SOA from cooking emissions.

Consistent with other studies, our work has shown that aldehydes are an important class of VOC precursors emitted from cooking emissions, and substantial efforts have been made to measure their emission factors depending on different cooking settings (heating temperature, cooking style, food, ingredients) (Klein et al., 2016a, 2016b). However, the contribution of aldehydes from cooking emissions is underrepresented in chemical transport models. Recently, McDonald et al. (2018) showed that the ambient concentrations of OA were underpredicted when aldehydes were not included in the box model calculations, suggesting that aldehydes, likely from food cooking, play an important role in atmospheric oxidation chemistry. Furthermore, Klein et al. (2019) showed that heavy polluters like restaurants play a significant role in contributing to the ambient cooking organic aerosol concentrations. In this study, we showed a large fraction of the SOA is derived from aldehyde precursors, with strong similarities in chemical composition. Therefore, it is important to consider the contribution of aldehyde chemistry in atmospheric models towards total OA budget. Furthermore, we demonstrated the importance of fragmentation reactions and their influence on OA properties such as volatility and chemical composition. Future work should therefore focus on measuring not only the SOA formation, but also the oxygenated VOCs formed due to fragmentation upon aging to provide insights into aging of cooking emissions.

Formation of SOA from cooking emissions in the atmosphere is likely influenced by emissions of POA, and other gas-phase precursors. Therefore, inclusion of POA during atmospheric processing of cooking emissions will likely influence the physicochemical properties of cooking SOA. For instance, with cooking POA being much less functionalized than SOA, inclusion of POA will likely decrease the system O:C (or $\overline{\mathrm{OSc}}$). However, POA from cooking emissions can undergo heterogeneous reactions in the atmosphere, thereby increasing O:C (or $\overline{\mathrm{OSc}}$). On the other hand, there could potentially be contributions from other gas-phase precursors or S/IVOCs emitted from cooking vapors that can result in SOA formation. These precursors can potentially contribute to SOA formation from cooking emissions, but their oxidative evolution in the atmosphere is not well understood.

Gas-particle partitioning of SOA can be further affected by non-ideal mixing, as well as morphology of the particles (Shiraiwa et al., 2013; Zuend and Seinfeld, 2012). Future work should investigate the effect of these parameters on cooking SOA properties and formation potential. To account for thermodynamic mixing favourability of the particles, Hansen solubility framework developed by Ye et al. (2016) can be implemented to provide insights into SOA mixing and yield enhancement. As shown in Ye et al. (2018) primary meat-cooking emissions can enhance SOA yield from α-pinene due to similarity in Hansen solubility parameters suggesting that primary meat cooking particles are miscible with α-pinene SOA. It should be noted that present study did not investigate the effect of atmospherically relevant seed particles as well as NO_x levels which are representative of typical urban environments. Since emissions upon entering the atmosphere gets mixed with background air, other source emissions, and diluted upon mixing thereby altering the gas-particle partitioning, and thus the total OA loading. Therefore, it is important to understand the changes in partitioning and miscibility of cooking emissions as the composition continually evolves with atmospheric processing. Additionally, as mentioned earlier cooking SOA undergoes large mass transfer limitations due to changes in the phase state of the SOA particles, making it more so important to experimentally determine the corresponding viscosity of cooking SOA. Therefore, future work should focus on measuring both the viscosity and miscibility of SOA derived from cooking emissions.

Data availability. The data are available upon request to the corresponding author.

Author Contributions. MT and AWHC designed research. MT collected and analyzed data. MT, YL and AWHC interpreted results. MT and AWHC wrote the manuscript with inputs from YL.

Competing interests. The authors declare that they have no conflict of interests.

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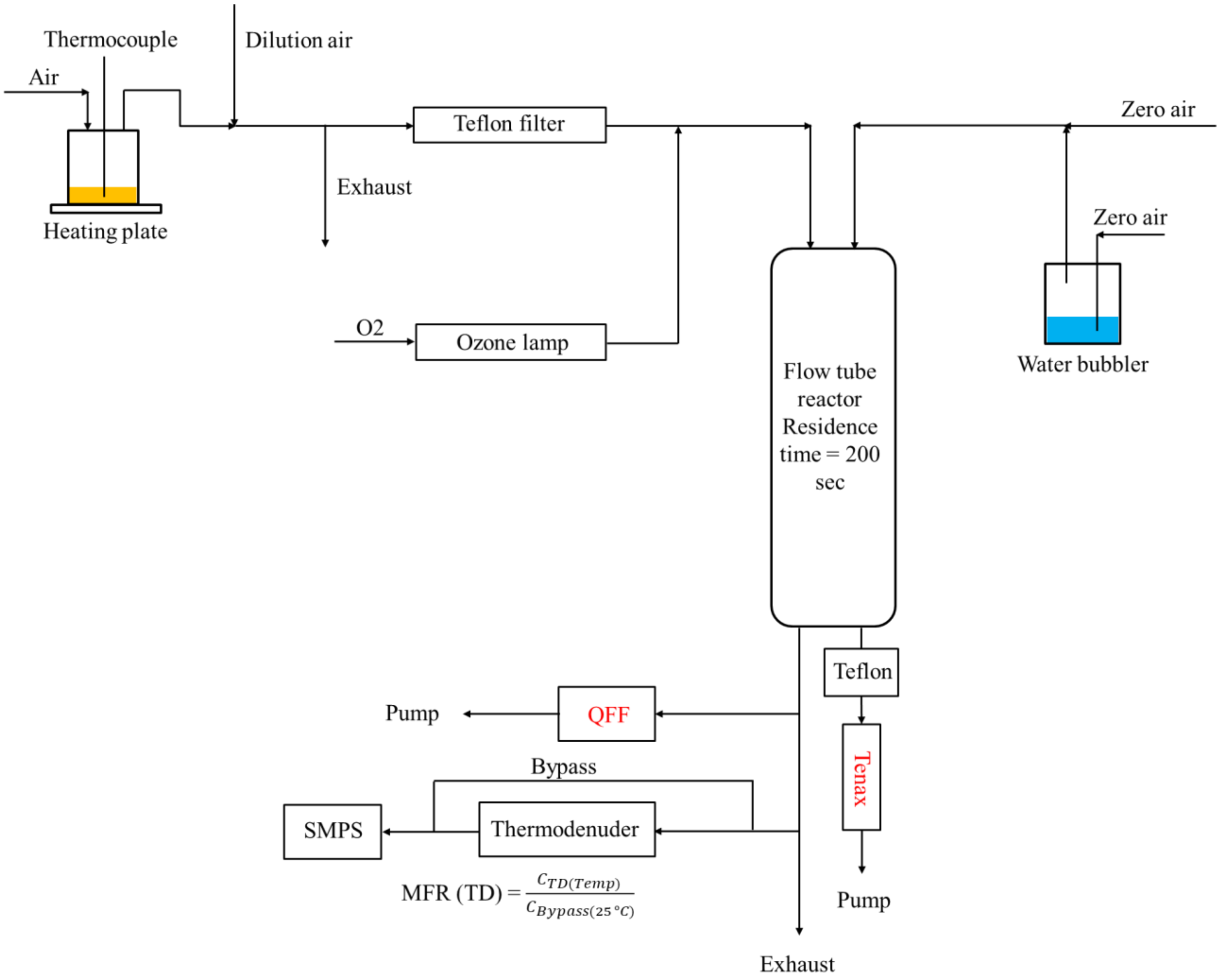


Figure 1. Experimental setup for oxidation of heated cooking oil emissions.

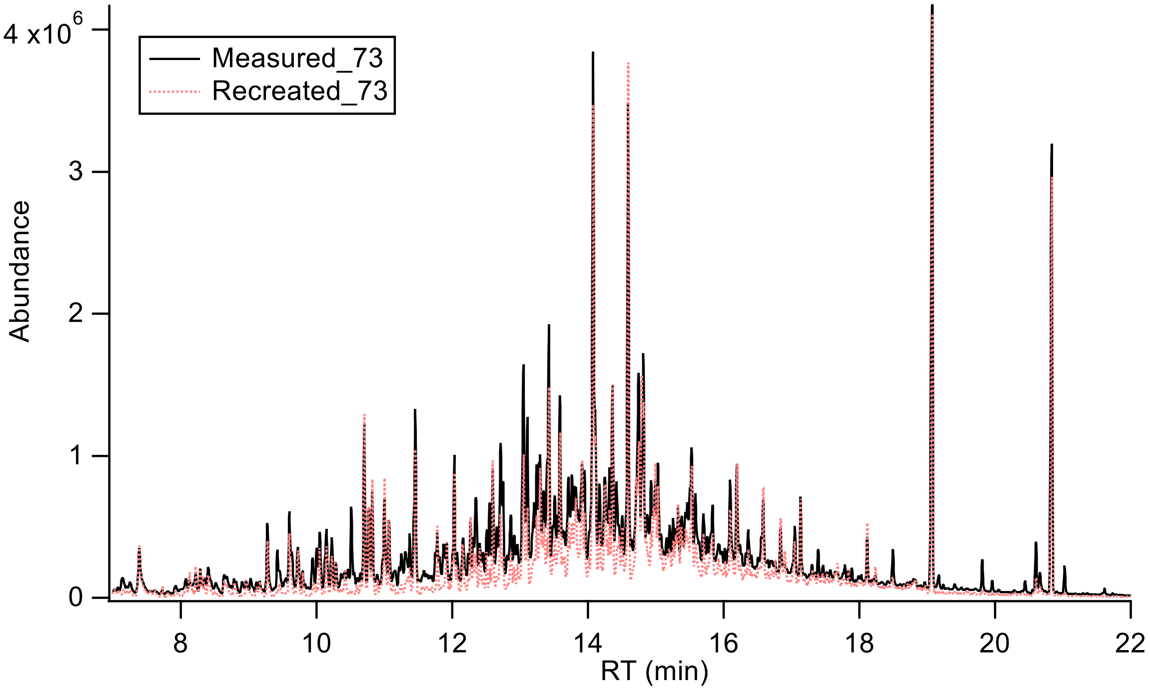


Figure 2. Highly complex mixture of canola oil SOA generated upon photooxidation. With known signal and mass fragmentation, signal of *m/z* 73 can be recreated based on pseudo parent ions (e.g. M-15 used in this study).

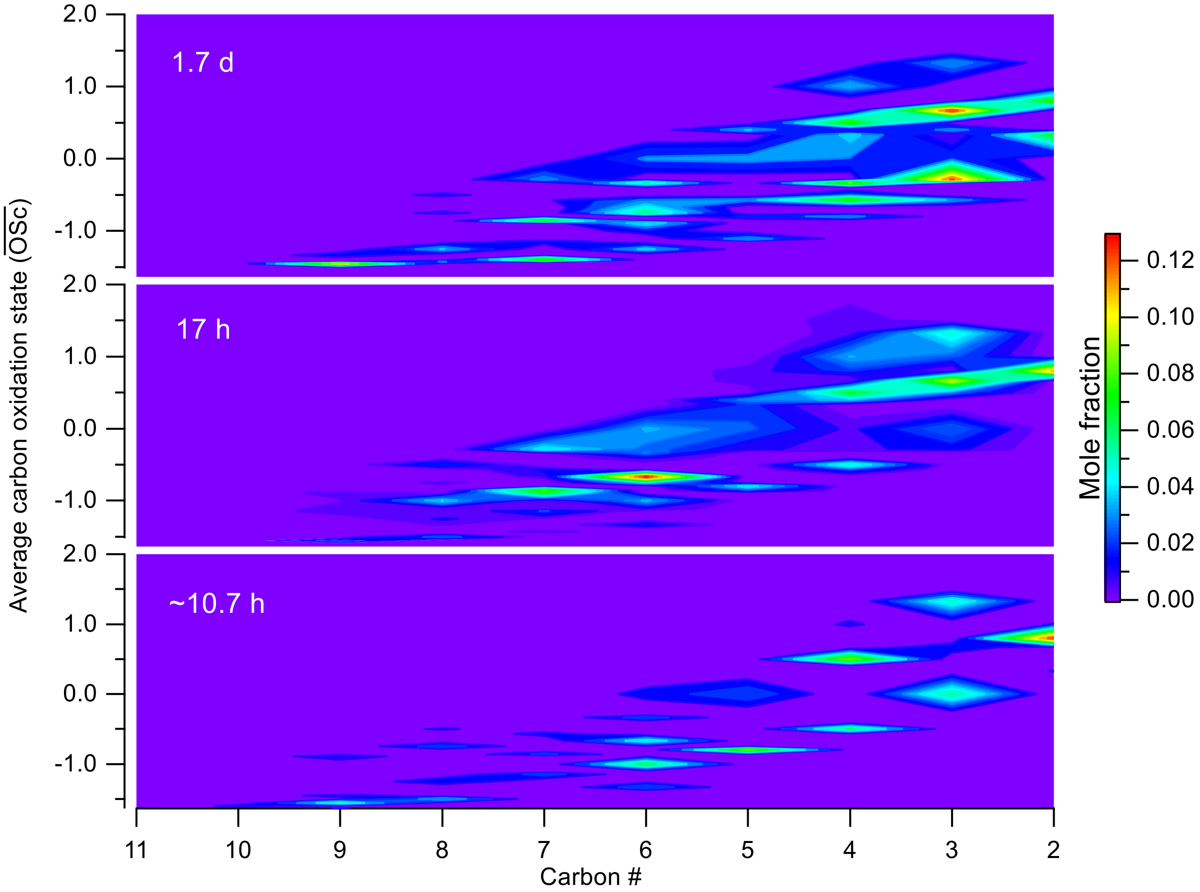


Figure 3. Evolution in $\overline{\text{OSc}}$-nc space for canola oil SOA under different conditions of photochemical aging. As the oxidation progresses in the atmosphere, more compounds are formed with smaller nc and higher $\overline{\text{OSc}}$ suggesting fragmentation to be a dominant pathway of oxidation for cooking emissions in the atmosphere.

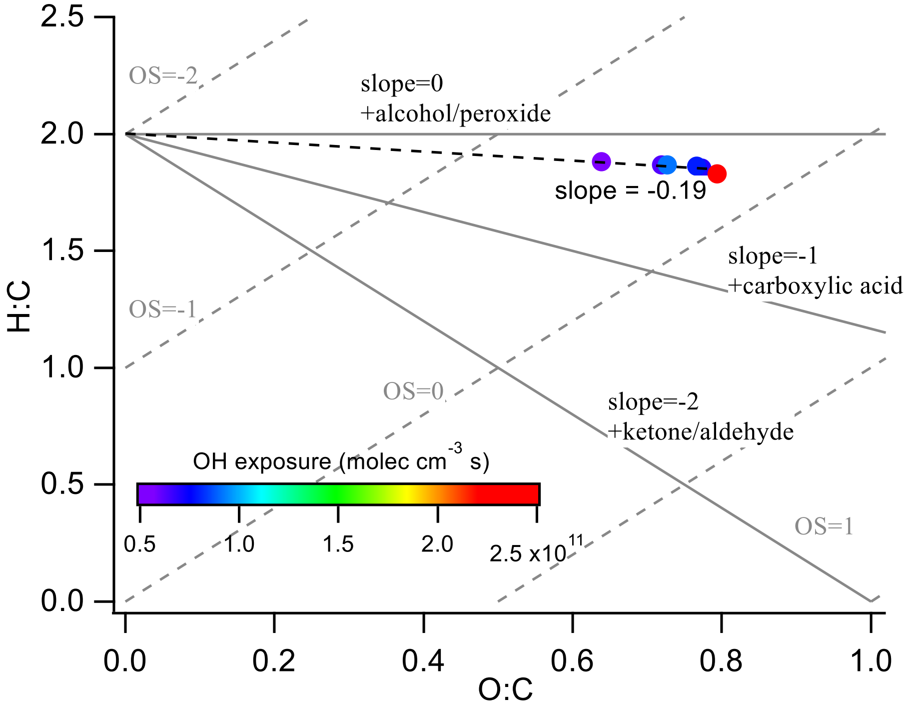


Figure 4. Van Krevelen diagram of canola oil SOA coloured by different OH exposure. In the background, average carbon oxidation state ($\overline{\mathrm{OSc}}$) and functionalization slopes are shown for reference. The slope of -0.19 for canola oil SOA corresponds to formation of both alcohol and carboxylic acid consistent with the chemical composition obtained from TD-GC/MS.

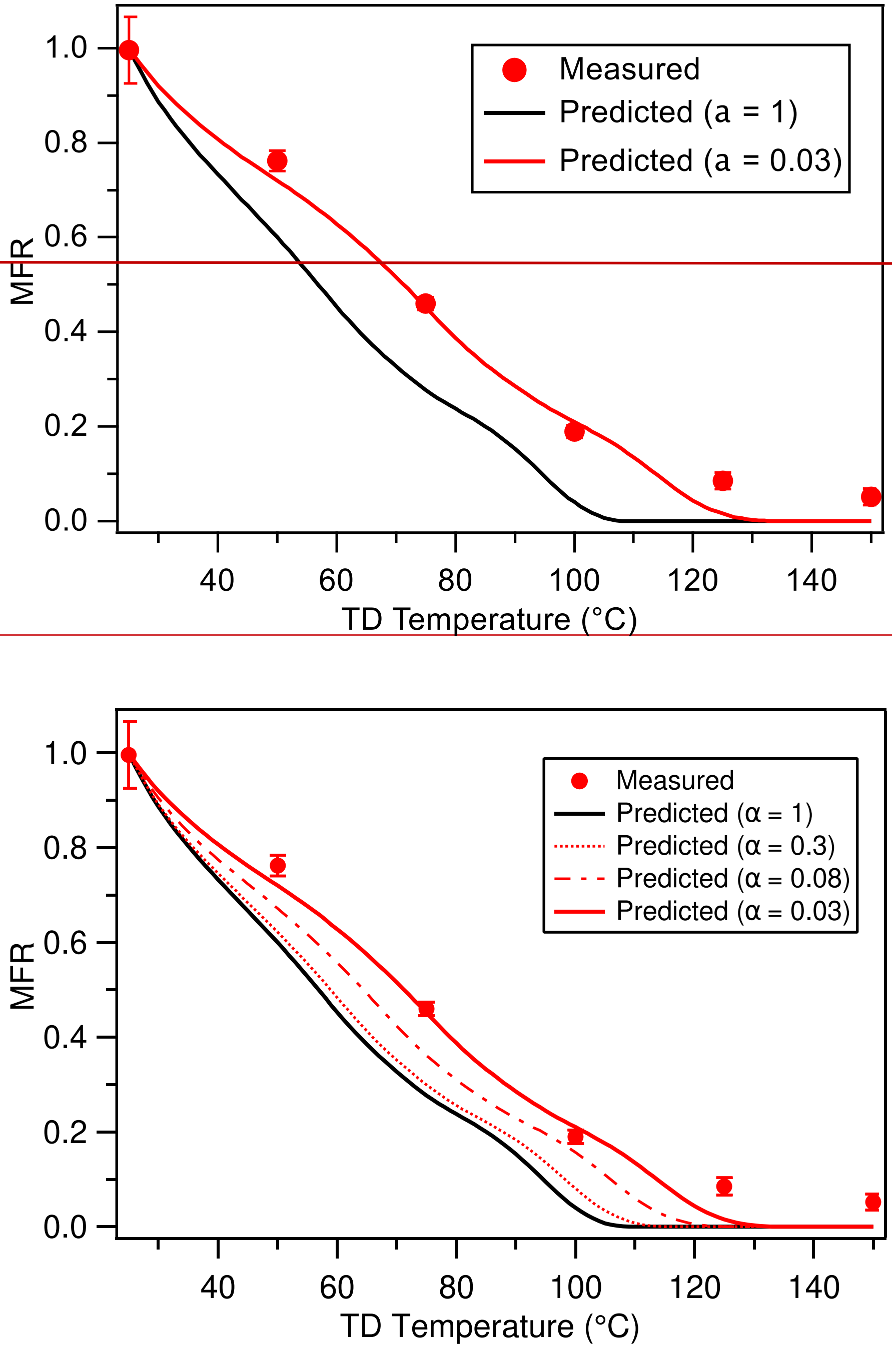
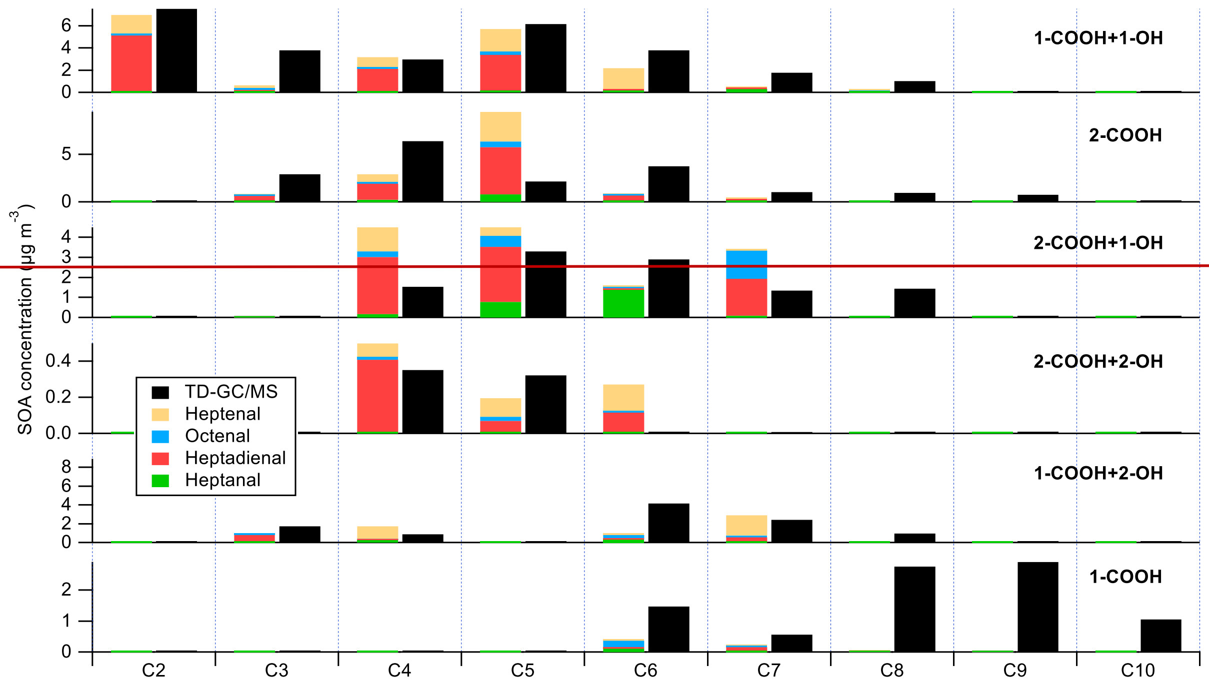


Figure 5. Mass thermogram of canola oil SOA at an OH exposure of 9.23×10^{10} molecules cm^{-3} s. The black line represents model simulations using $\alpha = 1$ underpredicting the measured MFR. The red line corresponds to model simulations using $\alpha = 0.03$ predicting the measurements reasonably well, therefore implying kinetic limitations in the system. The error bars represent ±1σ.



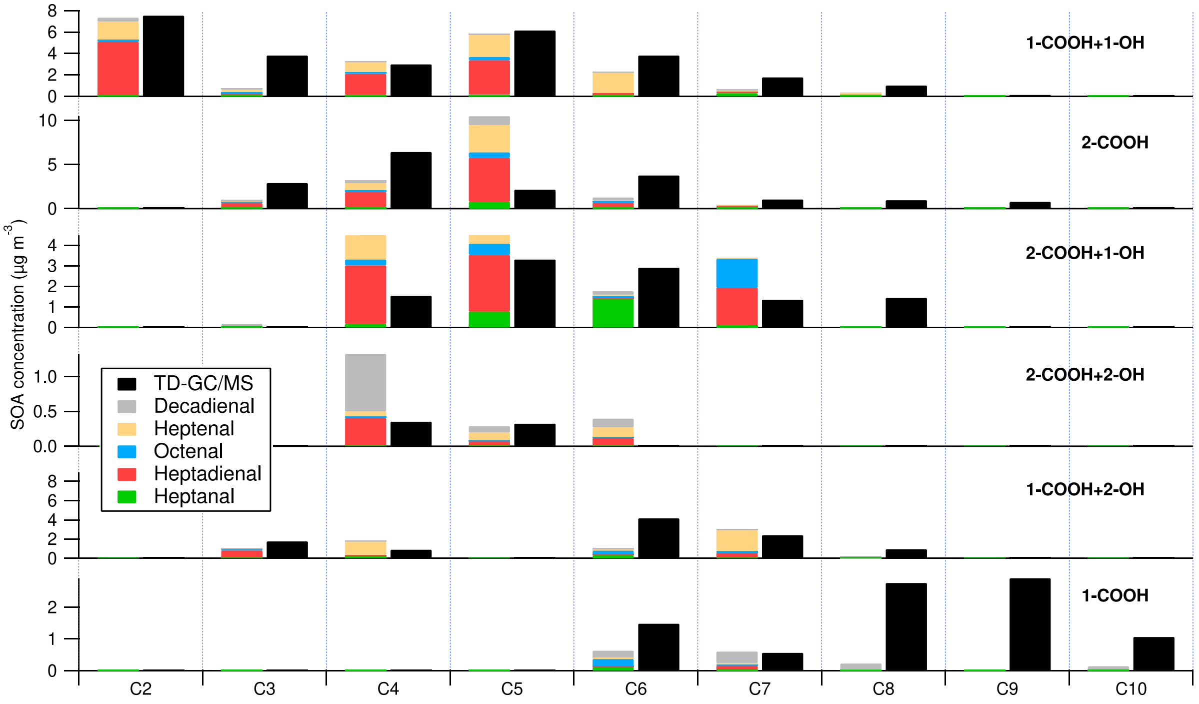


Figure 6. Prediction of different compounds formed at an OH exposure of 6.43×10^{10} molecules cm^{-3} s using product molar yields of heptanal, heptenal, octenal, ~~and~~ heptadienal, and decadienal. The total aldehydes products can explain the observed oil SOA products within a factor of half, while the inconsistency in prediction of some SOA products is likely caused by differences in gas-particle partitioning in both photooxidation systems.

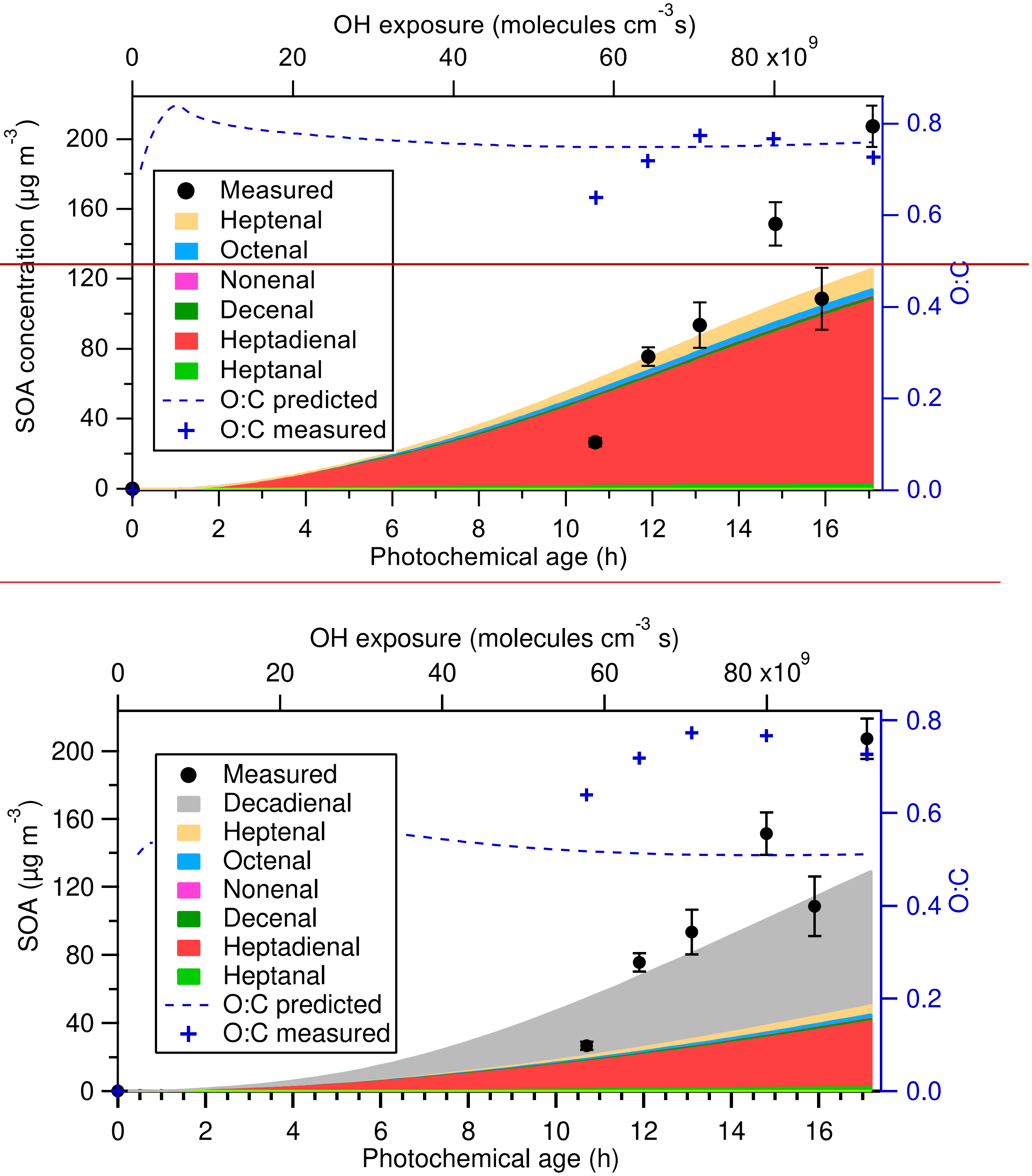


Figure 7. SOM prediction of SOA produced from different aldehydes with increasing photochemical age. The model overpredicts SOA formation at lower photochemical age, while underpredicts SOA formation by ~40% at higher photochemical age, suggesting that traditional VOC precursors cannot fully explain the SOA formation, and other gas-phase precursors maybe needed to better

constrain the formation of SOA at higher aging conditions. In addition, the SOM predicted O:C is within ~~±~~50~~20~~% of the measured O:C suggesting that the overall change in chemical composition of cooking SOA is predicted reasonably well.

Characterization of secondary organic aerosol from heated-cooking oil emissions: evolution in composition and volatility

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Table S1: Description of the experiments conducted in this study.

Exp.	Canola oil SOA ($\mu g\ m^{-3}$)	OH exposure (molecules cm^{-3} s)	Photochemical age (h)[a]
1	26.57±2.32	5.77×10^{10}	10.7
2	75.67±5.33	6.43×10^{10}	11.9
3	93.48±13.1	7.07×10^{10}	13.1
4	151.46±12.45	8.01×10^{10}	14.8
5	108.74±17.6	8.60×10^{10}	15.9
6	207.2±11.91	9.23×10^{10}	17.1
7	2670.4±170.85	2.20×10^{11}	40.7

a: calculated by assuming an average atmospheric OH concentration of 1.5×10^6 molecules cm^{-3} (Mao et al., 2009).

Table S2: List of all the model compounds used in this study to recreate 73 along with its $f_{M-15/73}$. SIMPOL.1 (Pankow and Asher, 2008) vapor pressure and its corresponding saturation concentrations are also listed for the compounds quantified in this study.

#	Carbon #	Molecular weight (MW)	Derivatized MW (M)	M-15	$f_{M-15/73}$ (NIST)	$f_{M-15/73}$ (MS detector response)	Vapor pressure (atm)	Saturation concentration ($\mu g\ m^{-3}$)
				2-COOH				
1	3	104	248	233	0.147	0.128±0.03	3.5E-07	29.5
2	4	118	262	247	0.24	0.273±0.06	1.3E-07	3.2[a]
3	5	132	276	261	0.67	0.254±0.1	4.9E-08	54.8[a]
4	6	146	290	275	0.198	0.143±0.04	1.9E-08	0.8[a]
5	7	160	304	289	0.187	0.198±0.07	7.0E-09	4.2[a]
6	8	174	318	303	0.147	0.182±0.08	2.6E-09	0.09[a]
7	9	188	332	317	0.268	0.313±0.22	9.9E-10	0.5[a]
				2-COOH + 1-OH				
8	3	120	336	321	0.029	0.062±0.014	2.3E-09	0.19
9	4	134	350	335	0.067	0.051±0.03	8.6E-10	0.02
10	5	148	364	349	0.146	0.141±0.05	3.2E-10	0.36
11	6	162	378	363	0.106[a1]	0.072±0.03	1.2E-10	5.3E-03
12	6	162	378	363	0.178[a2]	0.272±0.11	1.2E-10	5.3E-03
13	7	176	392	377	0.012	0.034±0.02	4.6E-11	2.7E-02

14	8	190	406	391	0.002	n/a	6.5E-12	6.4E-04
15	9	204	420	405	n/a	n/a	2.4E-12	3.2E-03
1-COOH + 2-OH								
16	3	106	322	307	0.04	0.058±0.01	4.9E-08	400.2
17	4	120	336	321	0.116	0.052±0.04	1.8E-08	150.6
18	5	134	350	335	n/a	n/a	6.9E-09	56.7
19	6	148	364	349	n/a	0.125±0.05	2.6E-09	21.3
20	7	162	378	363	n/a	0.019±0.01	9.8E-10	8.02
21	7	162	378	363	n/a	0.026±0.01	9.8E-10	8.02
22	7	162	378	363	n/a	0.068±0.01	9.8E-10	8.02
23	7	162	378	363	n/a	0.056±0.01	9.8E-10	8.02
24	8	176	392	377	n/a	0.023±0.02	3.7E-10	3.01
25	8	176	392	377	n/a	0.026±0.01	3.7E-10	3.01
26	9	190	406	391	n/a	0.033±0.02	1.4E-10	1.1
27	9	190	406	391	n/a	0.021±0.02	1.4E-10	1.1
2-COOH + 2-OH								
28	4	150	438	423	0.07	0.029±0.002	5.7E-12	1.4E-04
29	5	164	452	437	0.059	0.057±0.01	2.1E-12	2.4E-03
30	6	178	466	451	0.05	n/a	8.0E-13	3.5E-05
1-COOH + 1-OH								
31	2	76	220	205	0.123	0.11±0.01	1.9E-05	162375.7
32	3	90	234	219	0.479	0.48±0.15	7.5E-06	61085.0
33	4	104	248	233	0.024[b1]	0.053±0.02	2.8E-06	22979.9
34	4	104	248	233	0.226[b2]	0.099±0.05	2.8E-06	22979.9
35	4	104	248	233	0.179[b3]	0.147±0.06	2.8E-06	22979.9
36	5	118	262	247	0.043[c1]	0.034±0.01	1.1E-06	8644.9
37	5	118	262	247	0.06[c2]	0.064±0.01	1.1E-06	8644.9
38	5	118	262	247	0.341[c3]	0.245±0.12	1.1E-06	8644.9
39	6	132	276	261	0.04[d1]	0.043±0.01	3.9E-07	3252.2
40	6	132	276	261	0.089[d2]	0.068±0.02	3.9E-07	3252.2
41	6	132	276	261	0.215[d3]	0.106±0.06	3.9E-07	3252.2

42	7	146	290	275	0.06	0.071±0.01	1.5E-07	1223.5
43	8	160	304	289	0.109	0.05±0.008	5.6E-08	460.3
44	9	174	318	303	0.074	n/a	2.1E-08	173.1
1-COOH								
45	6	116	188	173	0.69	0.604±0.05	6.1E-05	5.7
46	7	130	202	187	1.18	0.656±0.02	2.3E-05	5.3
47	8	144	216	201	1.14	0.769±0.06	8.6E-06	4.8
48	9	158	230	215	1.24	0.815±0.04	3.2E-06	4.4
49	10	172	244	229	1.02	0.947±0.11	1.2E-06	3.9

a: Bilde et al. (2003).
a1: positional isomers
a2: positional isomers
b1: α-hydroxyisobutyric acid
b2: β-hydroxybutyric acid
b3: 3-hydroxybutyric acid
c1: α-hydroxyvaleric acid
c2: β-hydroxy-n-valeric acid
c3: 4-hydroxyvaleric acid
d1: 4-methyl 2-keto pentanoic acid
d2: 3-hydroxycaproic acid
d3: 5-hydroxyhexanoic acid

Table S3: List of all the model compounds used in this study to recreate 73 for single precursor oxidation experiments along with corresponding product molar yields.

M-15	**Heptanal + OH**	**2-heptenal + OH**	**2-octenal + OH**	**2,4-heptadienal + OH**	**2,4-decadienal + OH**
OH exposure (molec cm^{-3} s)[a]	7.71×10^{10}	6.2×10^{10}	6.02×10^{10}	5.34×10^{10}	5.72×10^{10}
OH reaction rate constant (cm^3 $molec^{-1}$ s^{-1})	3.0×10^{-11}[b]	4.4×10^{-11}[c]	4.1×10^{-11}[d]	4.6×10^{-10}[e]	4.6×10^{-10}[f]
1-COOH + 1-OH					
205	1.87E-07	7.38E-05	4.44E-05	9.40E-05	1.94E-05
219	6.85E-08	1.26E-05	1.29E-05	3.84E-06	1.74E-06
233	3.67E-06	2.97E-05	2.45E-05	2.66E-05	2.24E-06
247	4.49E-06	5.82E-05	5.03E-05	3.77E-05	4.52E-06
261	3.73E-06	4.78E-05	2.22E-05	0	9.07E-07

275	6.69E-06	4.35E-06	6.46E-06	0	0
289	0	5.70E-07	2.27E-05	0	7.30E-07
303	0	0	0	0	0
317	0	0	0	0	0
2-COOH					
233	1.96E-06	5.32E-06	9.26E-06	7.44E-06	4.05E-06
247	5.56E-06	2.35E-05	3.02E-05	1.98E-05	1.02E-05
261	1.82E-05	7.85E-05	9.24E-05	5.26E-05	3.0E-05
275	2.87E-06	5.00E-06	3.62E-06	5.20E-06	9.31E-06
289	3.19E-06	9.94E-07	1.34E-05	0	5.92E-07
303	0	0	0	0	0
317	0	0	0	0	0
2-COOH + 1-OH					
321	0	0	0	0	5.45E-06
335	3.11E-06	0	0	2.96E-05	1.76E-05
349	1.33E-05	3.02E-05	4.20E-05	2.61E-05	1.07E-05
363	2.29E-05	1.80E-05	7.28E-05	0	4.29E-06
377	1.63E-06	1.38E-06	1.72E-05	1.45E-05	1.46E-06
391	0	0	0.00016	0	0
2-COOH + 2-OH					
423	0	1.69E-06	1.90E-06	3.81E-06	2.18E-05
437	0	2.15E-06	2.73E-06	5.97E-07	2.28E-06
451	0	2.93E-06	0	9.19E-07	2.72E-06
1-COOH + 2-OH					
307	3.41E-07	3.32E-06	0	1.06E-05	4.95E-06
321	4.27E-06	3.82E-05	2.31E-05	0	5.09E-06
335	0	0	0	0	0
349	5.50E-06	0	6.69E-05	0	7.46E-06
363	0	4.42E-05	2.68E-05	4.86E-06	0
377	0	0	0	0	5.71E-06
391	0	0	0	0	0

		1-COOH			
173	3.70E-06	0	4.53E-05	0	8.99E-06
187	0	1.65E-06	2.64E-06	1.73E-06	1.09E-05
201	0	0	1.05E-06	0	6.0E-06
215	0	0	0	0	4.37E-07
229	0	0	0	0	3.33E-06

a: as calculated from the decay of cyclopentane in each set of experiment.

b: Atkinson and Arey (2003).

c: Davis et al. (2007).

d: Gao et al. (2009).

e: calculated by multiplying 2-heptenal OH reaction rate constant by a factor of 105 as obtained from Kwok & Atkinson (1995).

f: assumed same as 2,4-heptadienal.

Section 1. Sample calculation for aldehyde reaction timescales

The reaction rate constant of methacrolein is obtained from Atkinson and Arey (2003) and is as follows:

k_{OH} = 2.9E-11 cm^3 $molec^{-1}$ s^{-1}

k_{O3} = 1.2E-18 cm^3 $molec^{-1}$ s^{-1}

For lowest OH exposure at 5.77×10^{10} molecules cm^{-3} s, OH conc = 2.88×10^{8} molecules cm^{-3}, and ozone concentration was measured ~0.5 ppm which corresponds to 1.23×10^{13} molecules cm^{-3} assuming 1ppb O_3 = 2.46×10^{10} molecules cm^{-3}. At these photooxidation conditions, methacrolein reaction timescale is calculated as follows:

$$\tau_{OH} = \frac{1}{k_{OH}[OH]} = \sim 120\ s$$

$$\tau_{O_3} = \frac{1}{k_{O_3}[O_3]} = 1129\ min$$

For highest OH exposure at 2.2×10^{11} molecules cm^{-3} s, OH conc = 1.1×10^{9} molecules cm^{-3}, and ozone concentration was measured ~12.6 ppm which corresponds to 3.1×10^{14} molecules cm^{-3}. At these photooxidation conditions, methacrolein reaction timescale is calculated as:

$$\tau_{OH} = 31\ s$$

$$\tau_{O_3} = \sim 45\ min$$

Section ~~1~~2. Chemical characterization using TD-GC/MS

Gas-phase analysis: Tenax tubes were desorbed in the thermal desorption system (TDS) for thermal desorption with initial temperature at 50 °C held for 2 minutes followed by a ramp of 60 °C min^{-1} to 320 °C and held for 4 minutes. The analytes were transferred to the cooling injection system (CIS4, Gerstel) via a transfer line maintained at 300 °C during the run. The CIS4 was embedded with quartz wool filled quartz liner maintained at -40 °C during thermal desorption, and was heated to 320 °C at 12 °C s^{-1}, and held for 5 minutes at 320 °C. The GC column was held for 2 minutes at 40 °C and heated to 250 °C at a rate of 7 °C min^{-1} and held for additional 5 minutes at 250 °C.

Particle-phase analysis: 4 mm diameter filter punches were inserted into glass tubes (6 mm OD × 178 mm length, Gerstel) and placed in the TDS for thermal desorption. The temperature ramping program for thermal desorption was from 40 °C initial temperature held for 2 minutes followed by a ramp of 60 °C min^{-1} to 320 °C and held for 5 minutes at 320 °C. After the analytes were desorbed in the TDS, they were transferred to the cooling injection system (CIS4, Gerstel) via a transfer line maintained at 300 °C during the run. The CIS4 was embedded with quartz wool filled quartz liner maintained at 10 °C during thermal desorption to preconcentrate the desorbed analytes. The CIS was heated from 10 °C to 320 °C at 12 °C s^{-1}, and held for an additional 7 minutes at 320 °C. The GC column was heated from 40 °C to 300 °C at a ramp of 10 °C min^{-1} and held for 5 minutes at 300 °C. All samples were analyzed under electron impact at 70 eV using a standard tungsten filament with a source temperature at 230 °C. The MS was operated at 3.1 scans s^{-1} with an acquisition range from mass-to-charge (*m/z*) ratio 35 to 500.

Section ~~2~~3. Procedure to recreate *m/z* 73

Step 1: Extract all model M-15 ions from the total ion chromatogram.

Step 2: Divide each M-15 ion with its corresponding $f_{M-15,73}$. Wherever, NIST $f_{M-15,73}$ was not available, $f_{M-15,73}$ was calculated from the instrument detector response.

Step 3: Since higher *m/z* ions are susceptible to fragmentation under high electron ionization efficiency (70 eV in our study), therefore caution must be taken in recreating M-15 ions so as to avoid double counting of actual (or real) peaks. An example of this scenario is shown in Fig. S4 (a), using an example chromatogram of *m/z* 233, 335 and 349. As shown in Fig. S4 (a), *m/z* 233 has large number of fragments from higher *m/z* ions, and some of these fragments belong to actual *m/z*. Peaks at retention time (RT) = 13.059 min corresponds to *m/z* 335 while at RT = 14.599 min belongs to *m/z* 349. Therefore, in order not to overestimate these peaks, fragments of higher *m/z* should be set to zero when recreating smaller *m/z*.

Step 4: Repeat step 3 iteratively for remaining M-15 ions to minimize the effect of double counting, and only accounting for signal from actual or real peaks as shown in Fig. S4 (b).

Step 5: Add all M-15 together to recreate 73 as shown in Fig. 2 (main text) with scatter plot shown in Fig. S4 (c).

Section ~~3~~4. Sample calculation for Sect. 3.3.1 (formation of particle-phase oxidation products)

Estimation of SOA formation potential using product yields of VOC precursor oxidation products were done as follows:

M-15 ion = 219 corresponds to 3-hydroxypropanoic acid has a yield of 1.26E-05 upon photooxidation of 2-heptenal which is used to estimate SOA formation for canola oil photooxidation using Eq. (4) in main text.

Where, γ_{ij} represents yields of products (*i*) from photooxidation of 2-heptenal (*j*) (obtained from Table S3), and ΔVOC_j represents the decay in concentration of 2-heptenal during photooxidation of canola oil vapors, and was calculated based on the measured OH exposure.

Therefore, for 3-hydroxypropanoic acid (or M-15 = 219) the SOA formation can be predicted as:

$SOA_{pred} = \quad 1.26\mathrm{E}-05 * 299 * 90 = 0.339$ µg m^{-3}.

Similarly, using above methodology SOA formation can be estimated from other VOC precursors such as heptanal, 2-octenal, ~~and~~ 2,4-heptadienal, and 2,4-decadienal.

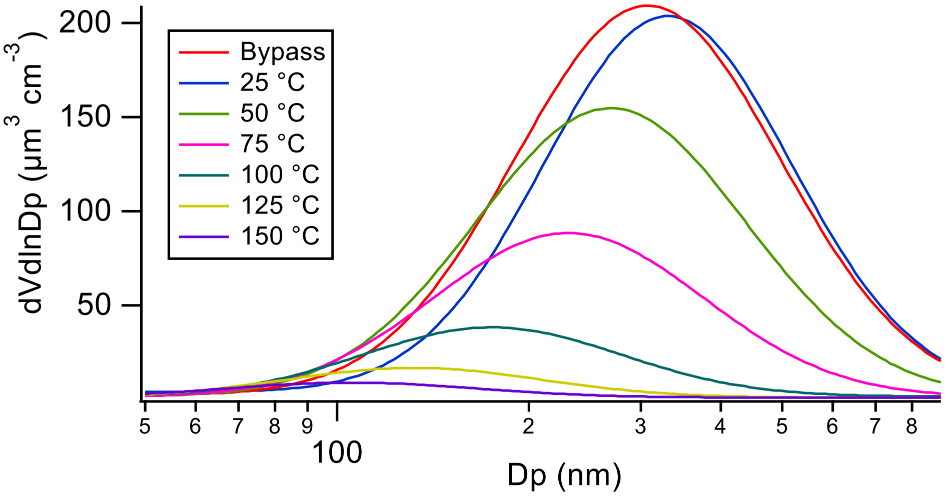


Figure S1. Particle volume distribution of canola oil SOA at an OH exposure of 9.23×10^{10} molecules cm^{-3} s measured by SMPS when subject to heating in a thermodenuder. The volume mode diameter shifts from 332 nm at 25 °C to 106 nm at 150 °C corresponding to a decrease in volume concentration of ~96%.

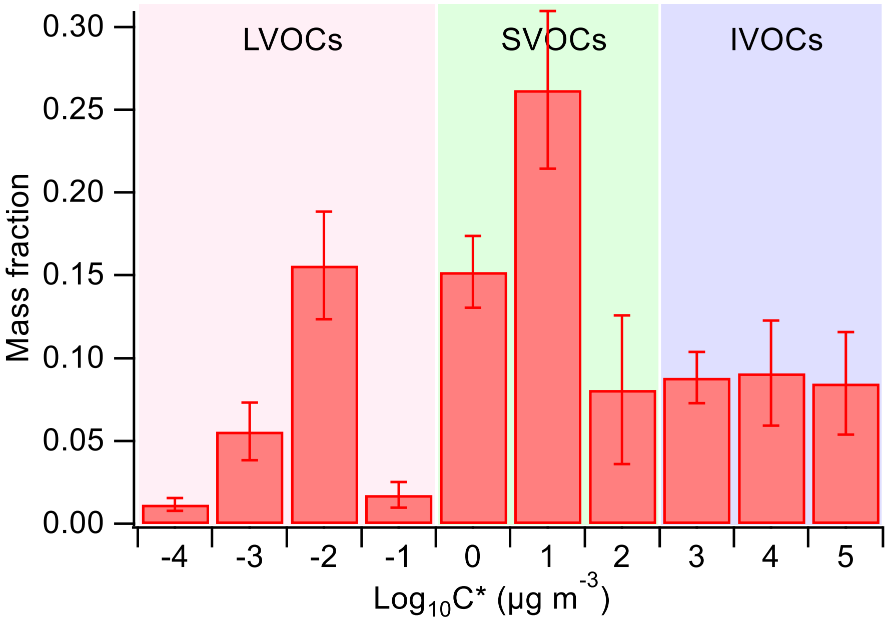


Figure S2. Volatility distribution of canola oil SOA at an OH exposure of 9.23×10^{10} molecules cm^{-3} s. The volatility distribution corresponds to 24% mass in LVOCs, ~50% in SVOCs, and 26% in IVOCs. The error bars represent ±1σ.

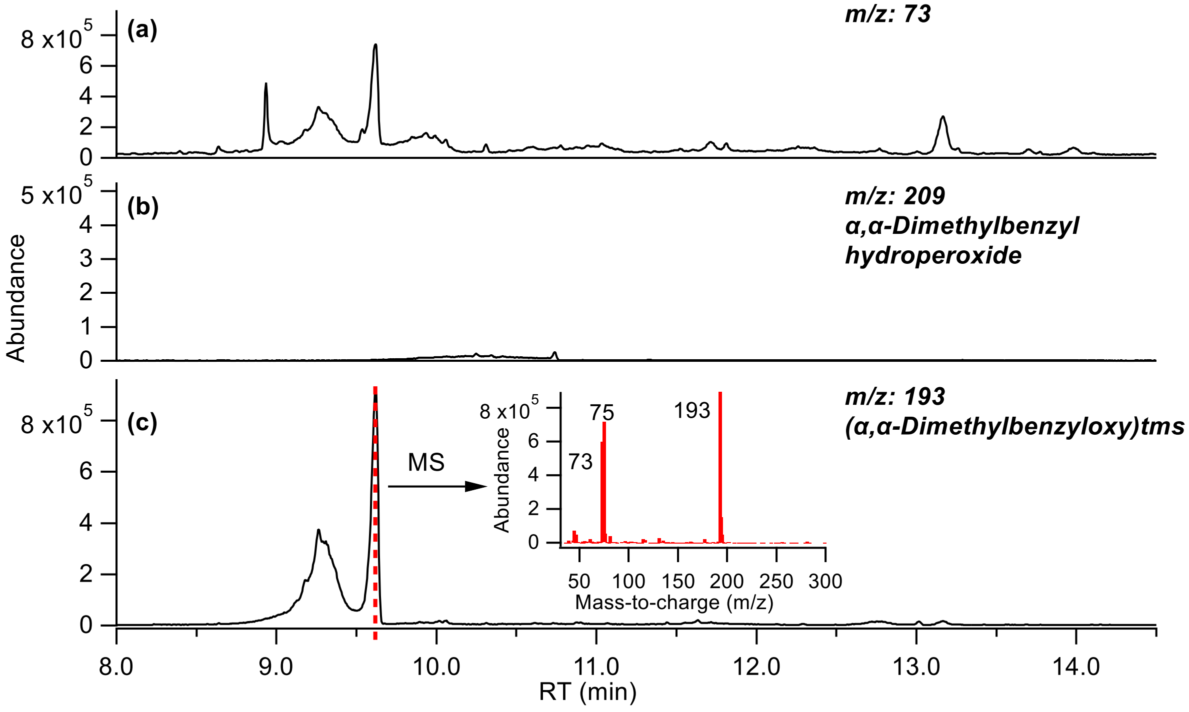
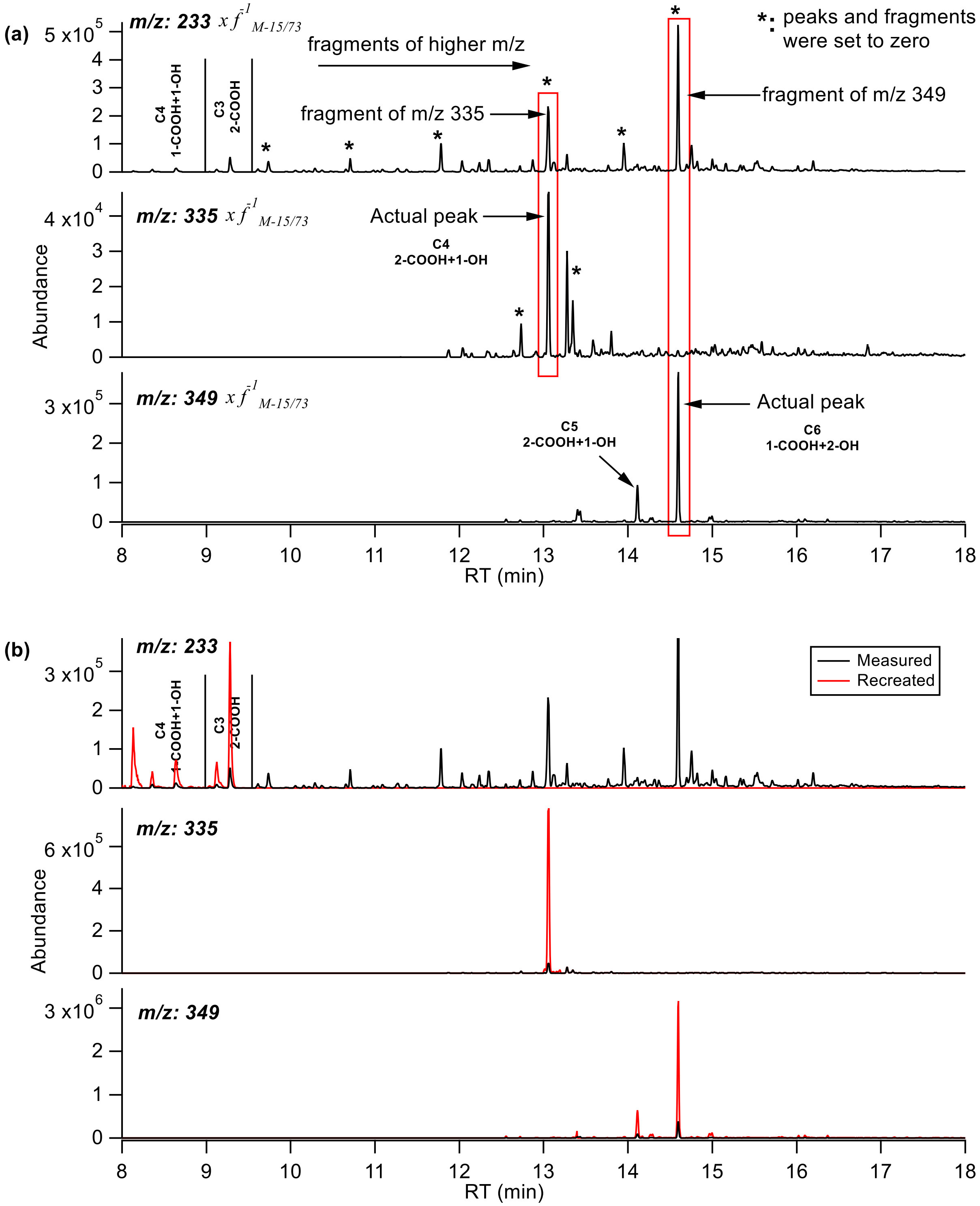
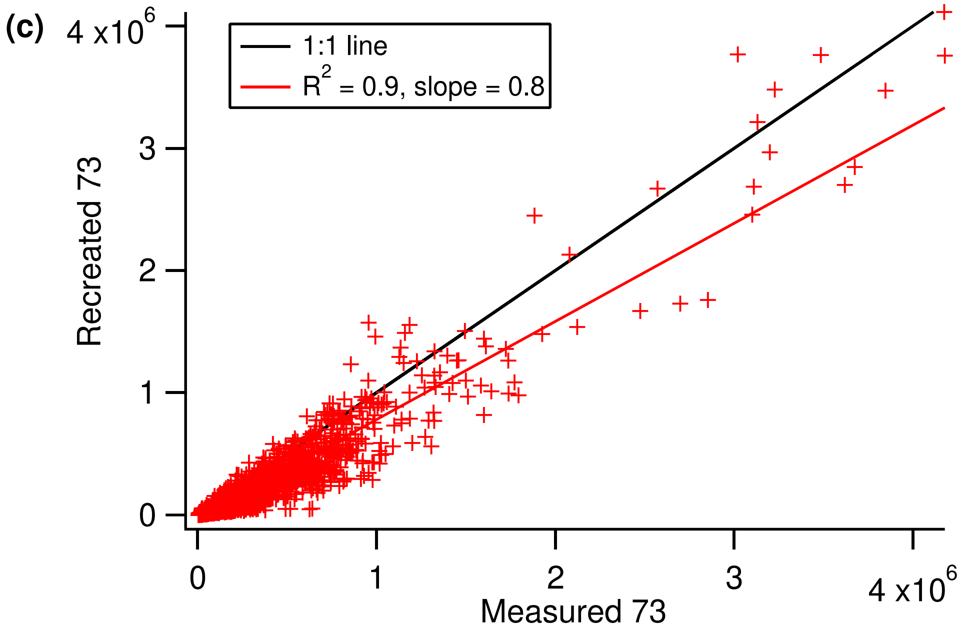


Figure S3. Chromatogram of cumene hydroperoxide upon *in situ* derivatization (a). Based on the analytical technique discussed in main text, extracted M-15 (m/z = 209) chromatogram of cumene hydroperoxide contains no peaks as shown in panel (b), instead the derivatized form of R-OH is observed as shown in panel (c) along with its mass spectrum shown in the inset.





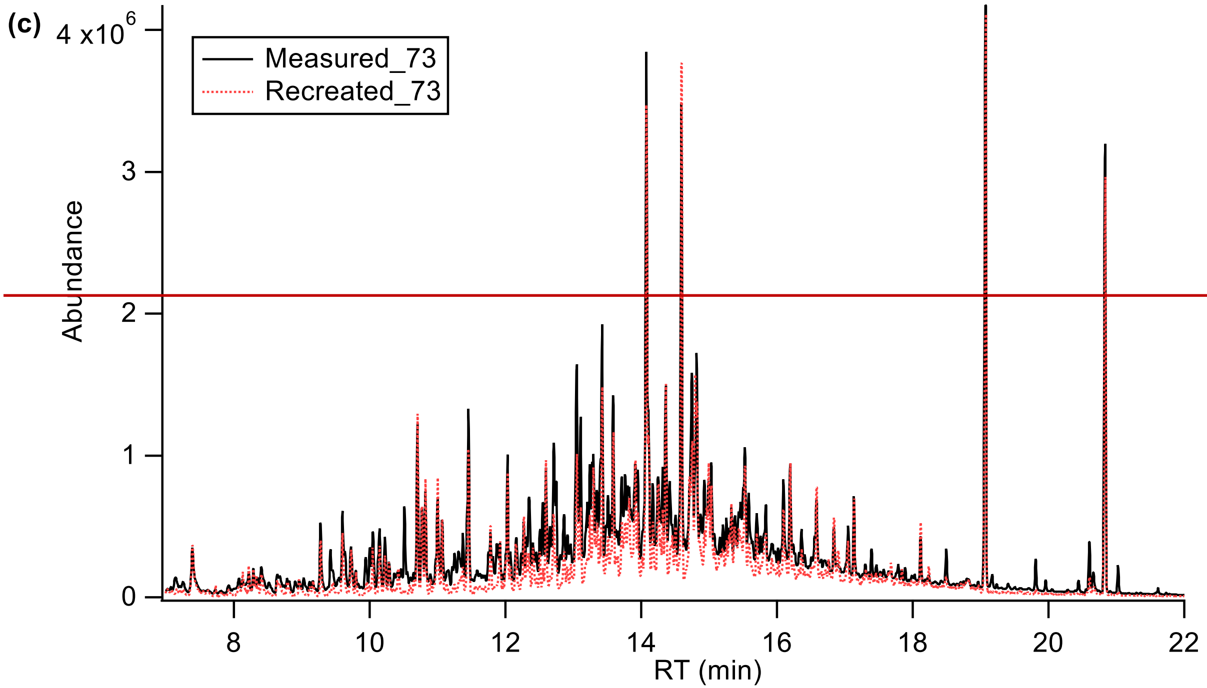
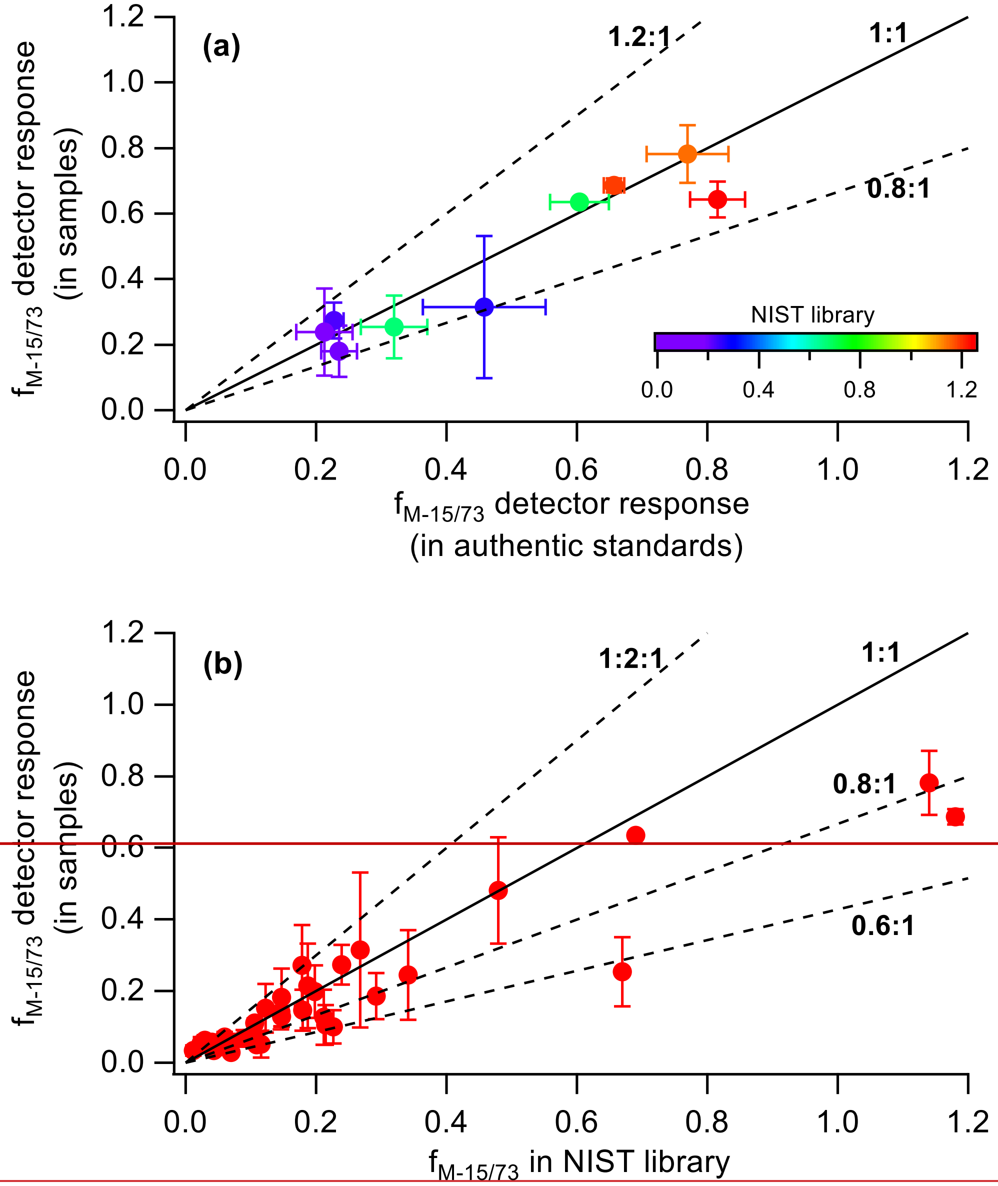


Figure S4. Illustration to recreate *m/z* 73. (a) shows the unprocessed M-15 chromatograms obtained from canola oil photooxidation highlighting that lower *m/z* ions are susceptible to interference from higher *m/z*, therefore appropriate processing (refer to Sect. ~~2~~3, step 3) of chromatograms should be carried out to account for these interferences. (b) chromatograms obtained after cleaning of fragments from each model *m/z*. (c) total ion chromatogram of measured and recreated 73.


(a)
1.2:1
1:1
0.8:1
NIST library
0.0
0.4
0.8
1.2
$f_{M-15/73}$ detector response (in samples)
$f_{M-15/73}$ detector response (in authentic standards)
0.0
0.2
0.4
0.6
0.8
1.0
1.2
(b)
1:2:1
1:1
0.8:1
0.6:1
$f_{M-15/73}$ detector response (in samples)
$f_{M-15/73}$ in NIST library
0.0
0.2
0.4
0.6
0.8
1.0
1.2

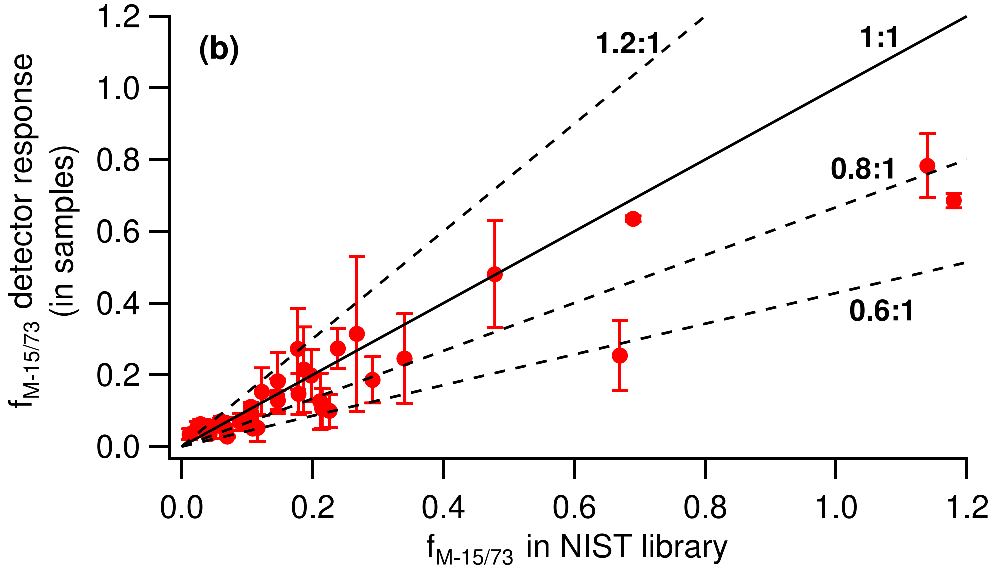


Figure S5. Calculated uncertainties in pseudo molecular ion fraction of model components. (a) compares $f_{M-15/73}$ from instrument detector response in samples vs that in authentic standards colour coded with $f_{M-15/73}$ available in NIST library. (b) shows the comparison of $f_{M-15/73}$ obtained from instrument detector response to that available in NIST library. Both comparisons show that the uncertainty in measurement of $f_{M-15/73}$ is within 20% for measured compounds except for tartaric acid which is within 40% uncertainty.

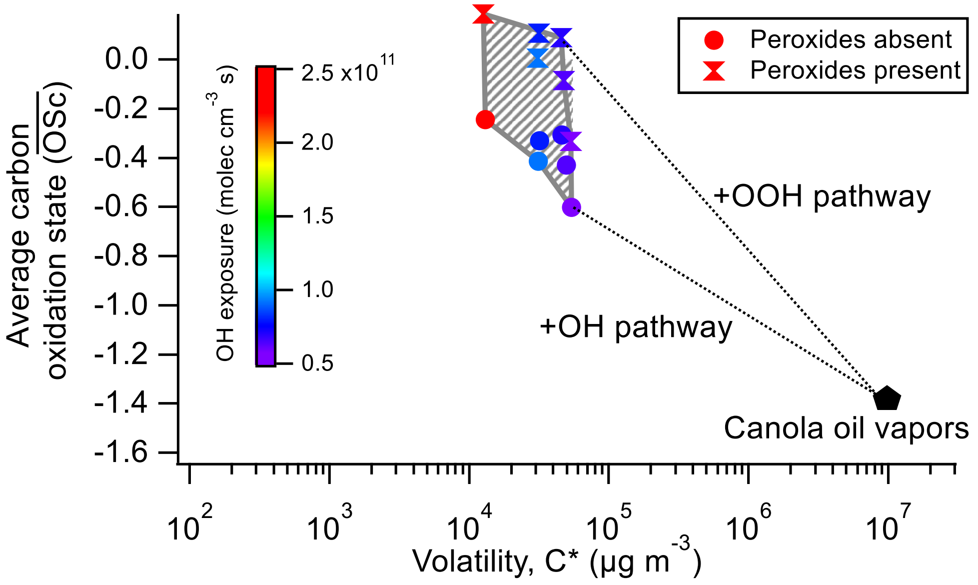


Figure S6. 2D-VBS for canola oil SOA upon photochemical aging in the atmosphere. The coloured markers represent bulk volatility of SOA under different photochemical aging conditions, while the black marker represents properties of canola oil vapors before oxidation. The shaded area corresponds to formation of SOA products with the uncertainty associated in identifying hydroxyl and peroxide groups. If all hydroxyl groups were instead classified as peroxide groups, the $\overline{OSc}$ increases but the bulk volatility of SOA shows a minor decrease suggesting that classification of peroxide groups as hydroxyl groups has little effect on estimation of volatility.

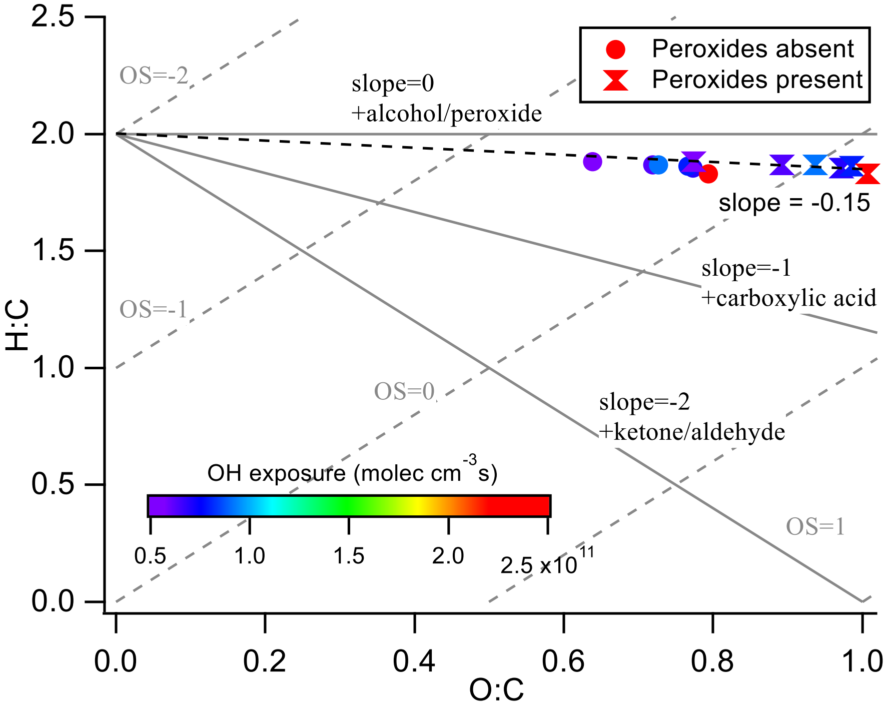
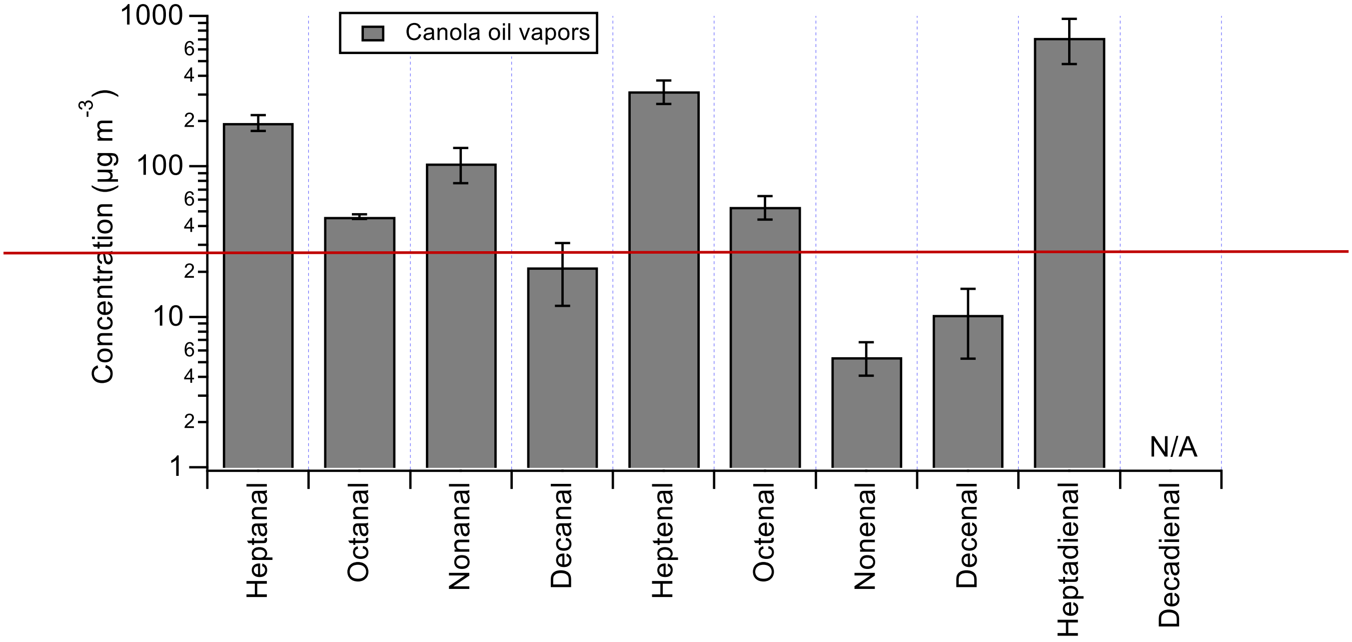


Figure S7. Van Krevelen diagram of canola oil SOA coloured by different OH exposure considering the presence of peroxides. The slope of -0.15 is observed when the formation of peroxides is considered in canola oil SOA similar to that of no-peroxides assumption (Fig. 4, main text).



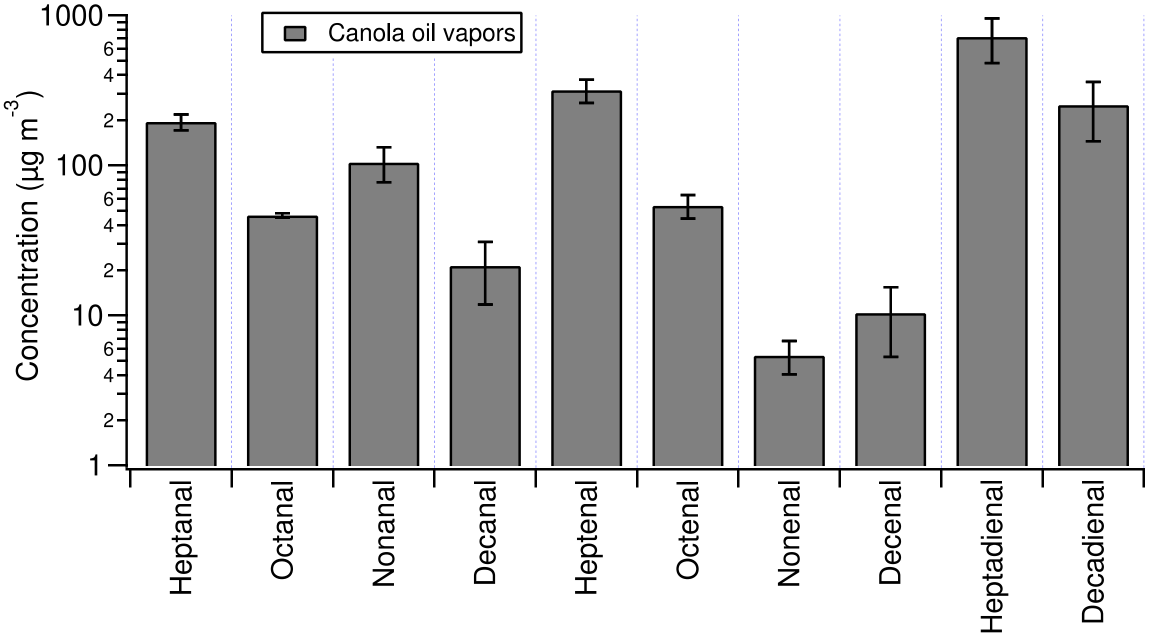
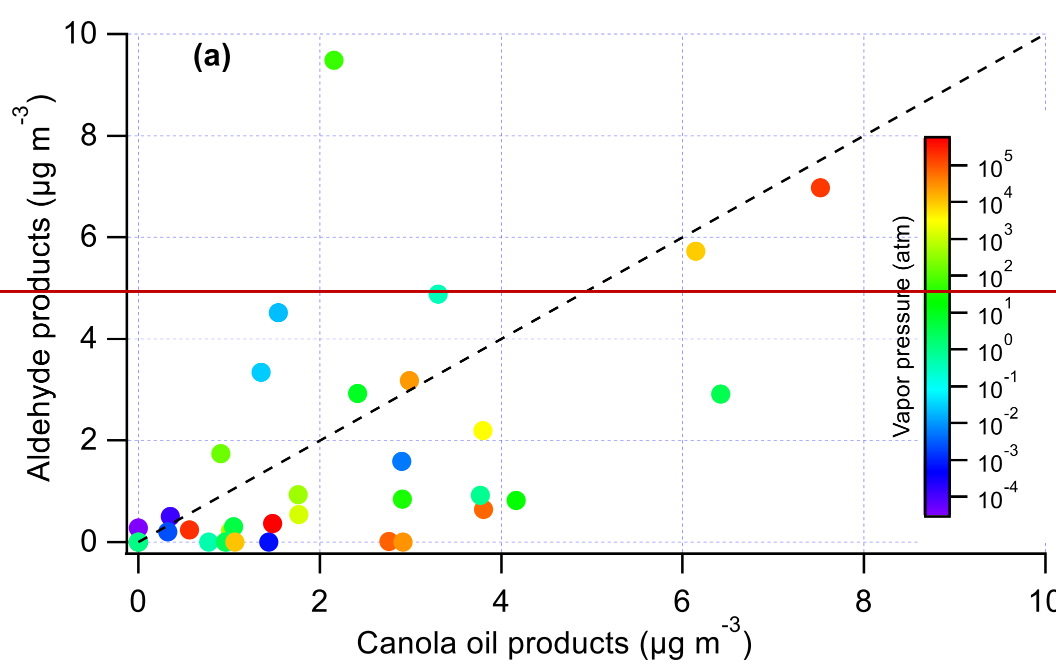
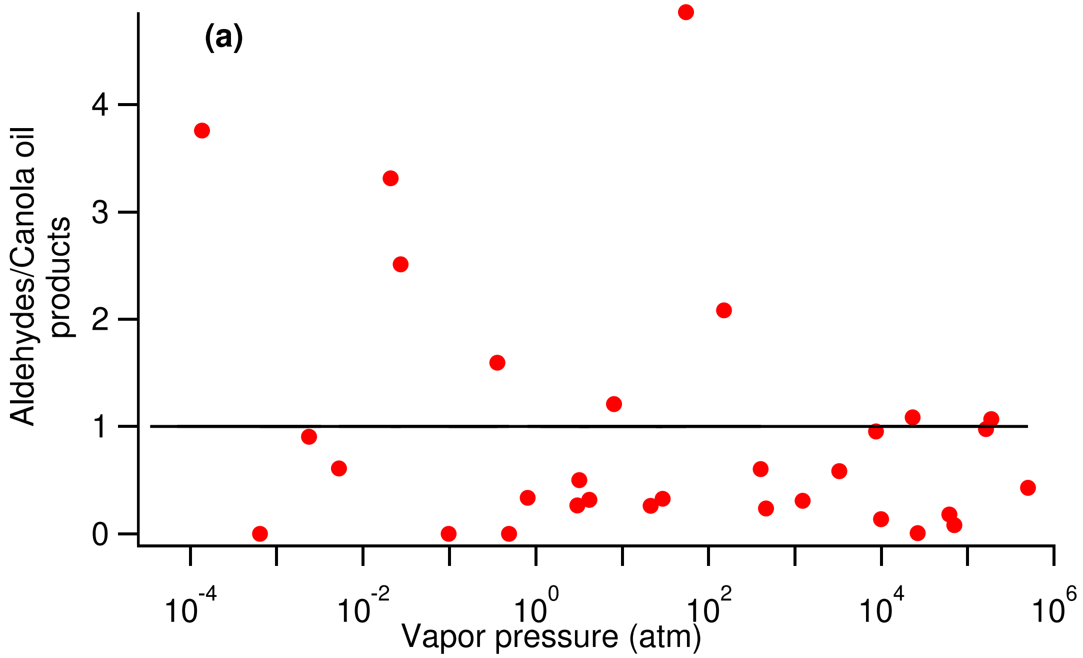
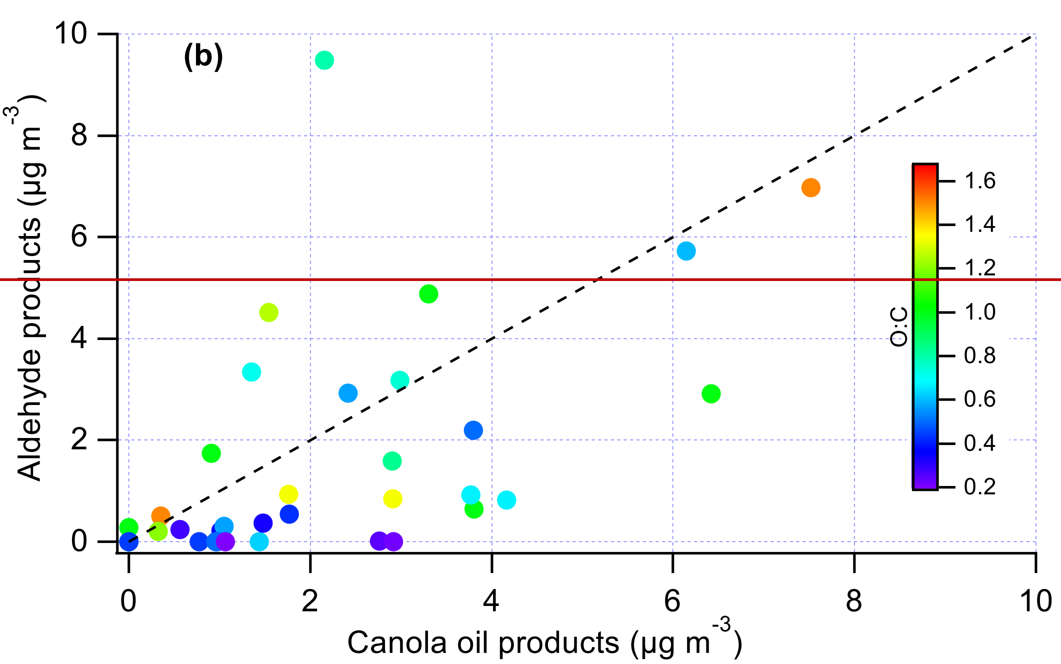
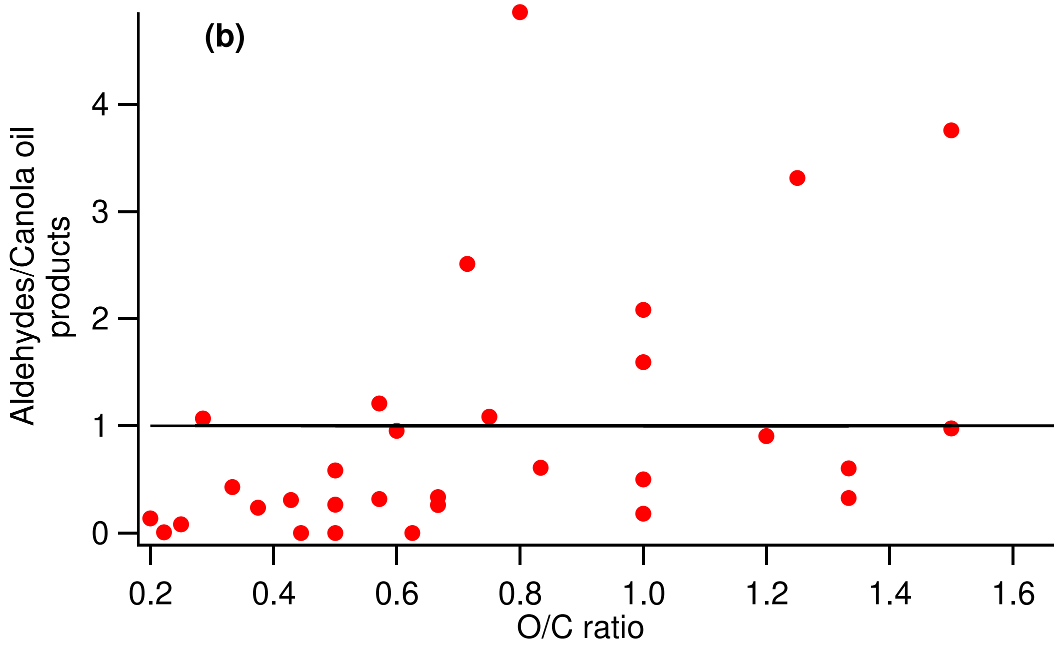


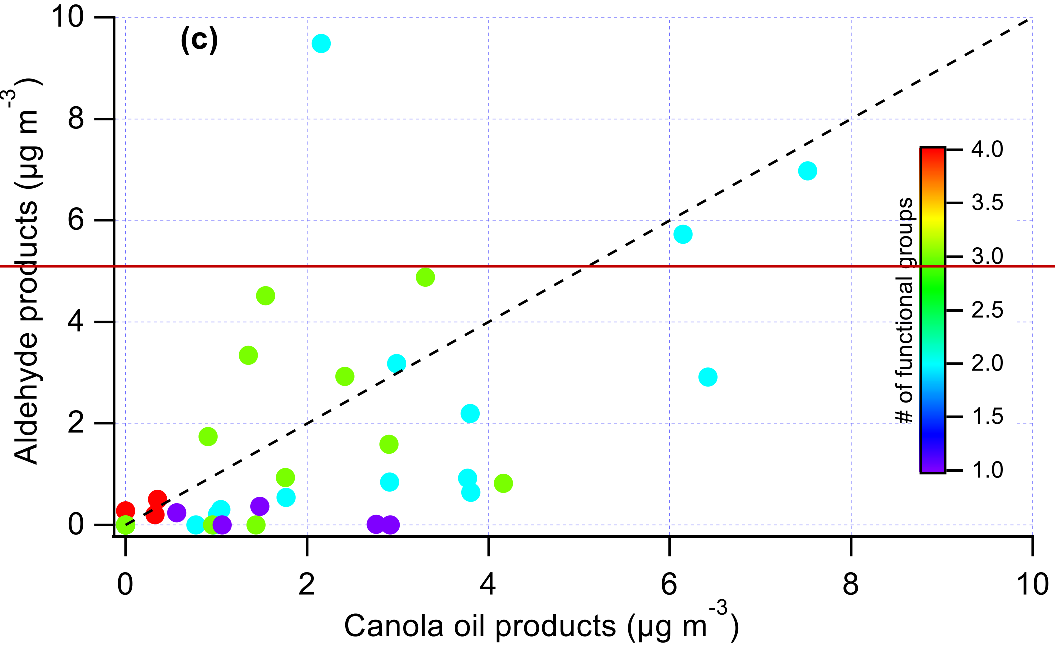
Figure S8. Concentration of different aldehydes quantified in this study.

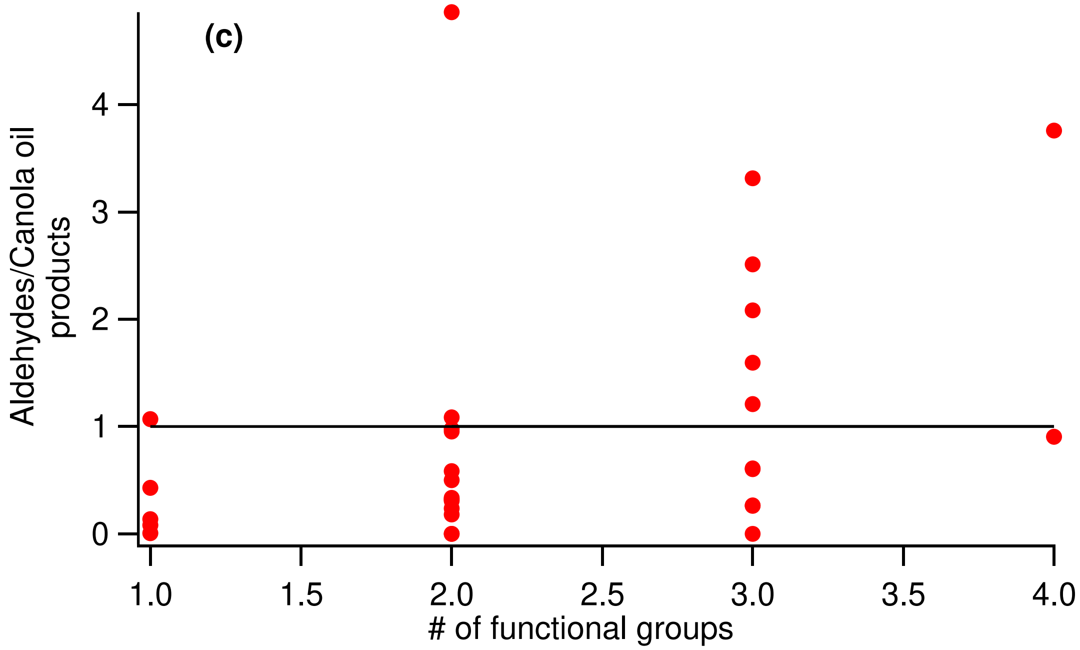

(a)
Aldehyde products (µg m^{-3})
Canola oil products (µg m^{-3})
0
2
4
6
8
10
Vapor pressure (atm)
10^{-4}
10^{-3}
10^{-2}
10^{-1}
10^{0}
10^{1}
10^{2}
10^{3}
10^{4}
10^{5}

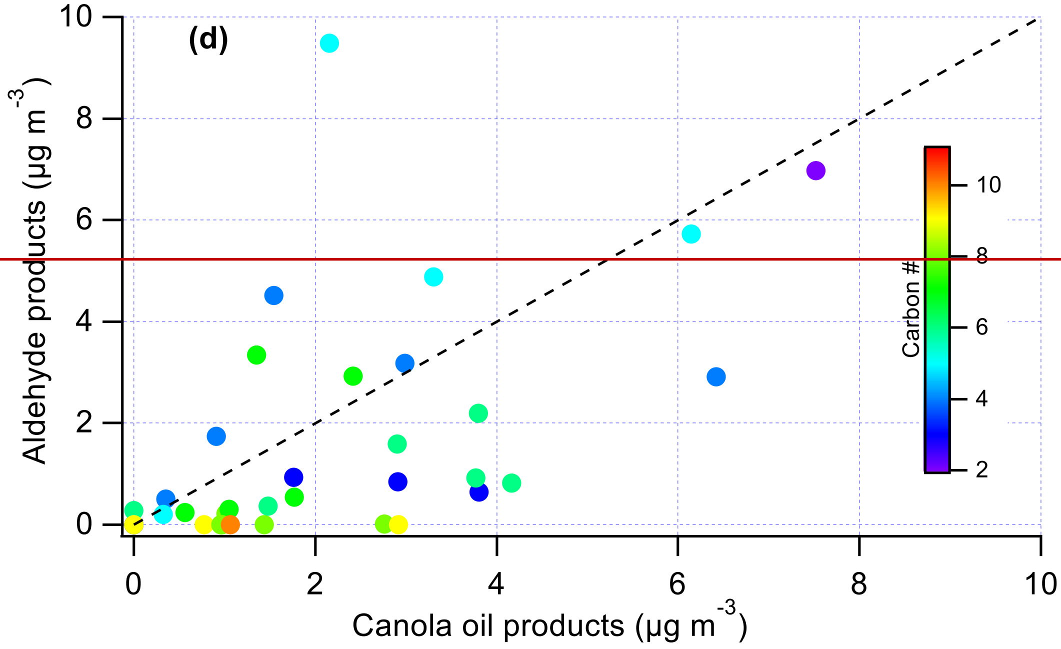

(a)
Aldehydes/Canola oil products
0
1
2
3
4
10^{-4}
10^{-2}
10^{0}
10^{2}
10^{4}
10^{6}
Vapor pressure (atm)


(b)
Aldehyde products (µg m^{-3})
Canola oil products (µg m^{-3})
O:C
0
2
4
6
8
10
0
2
4
6
8
10
1.6
1.4
1.2
1.0
0.8
0.6
0.4
0.2


(b)
Aldehydes/Canola oil products
O/C ratio
0
1
2
3
4
0.2
0.4
0.6
0.8
1.0
1.2
1.4
1.6


(c)
Aldehyde products (µg m^{-3})
Canola oil products (µg m^{-3})
# of functional groups
0
2
4
6
8
10
1.0
1.5
2.0
2.5
3.0
3.5
4.0


(c)
Aldehydes/Canola oil products
# of functional groups
0
1
2
3
4
1.0
1.5
2.0
2.5
3.0
3.5
4.0



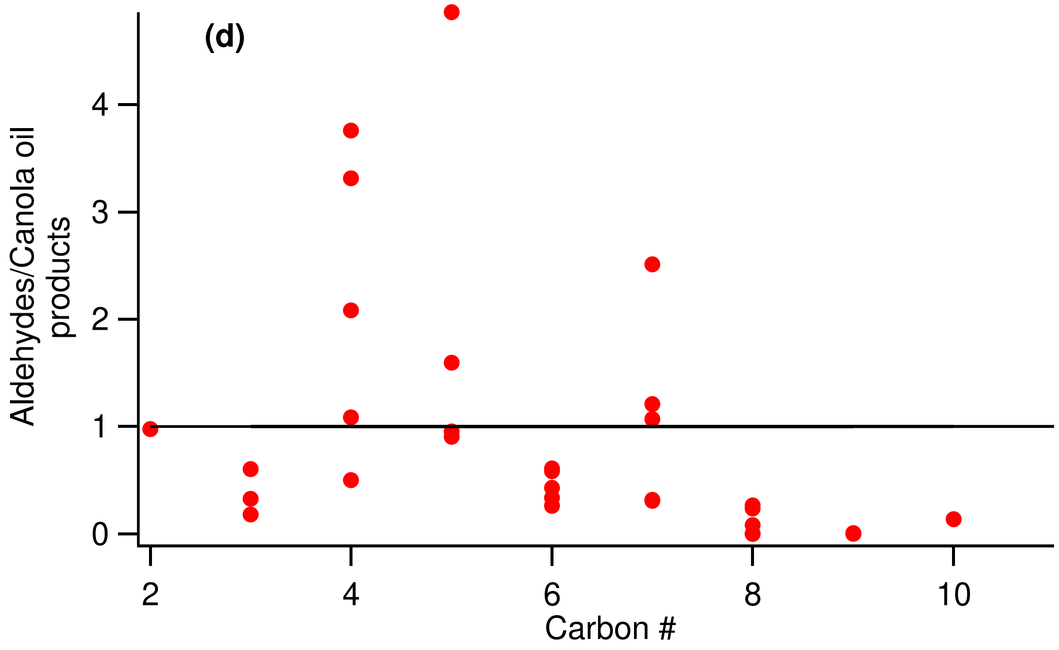
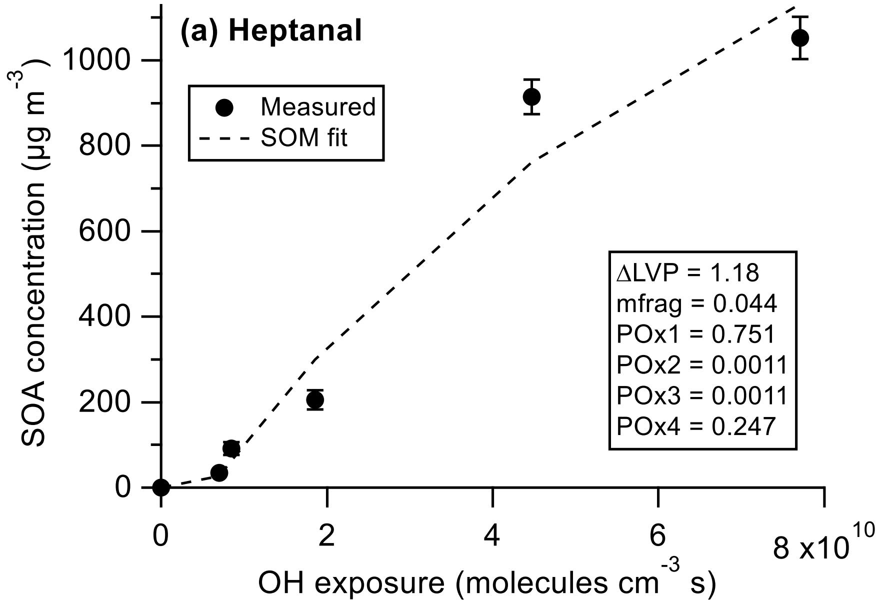
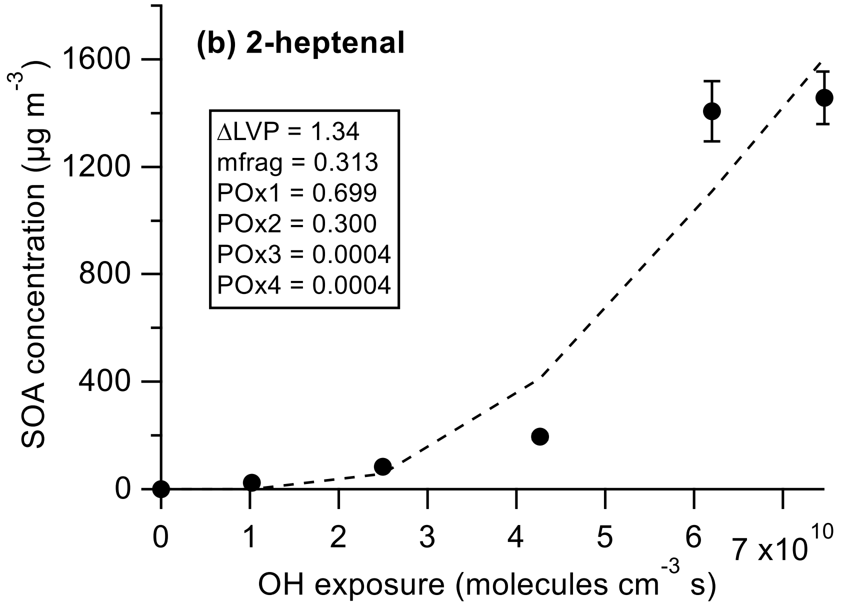


Figure S9. Comparison of the ratio of GC measured products from aldehydes photooxidation to canola oil photooxidation ~~and aldehydes photooxidation~~ by vapor pressure (a), O/C ratios (b), # of functional groups (c), and carbon # (d). In general, aldehydes SOA products are underestimated for lower O/C ratios and number of functional groups, suggesting that canola oil SOA favors partitioning of more oxygenated compounds than SOA formed from aldehydes.



(a) Heptanal
Measured
SOM fit
ΔLVP = 1.18
mfrag = 0.044
POx1 = 0.751
POx2 = 0.0011
POx3 = 0.0011
POx4 = 0.247
SOA concentration (μg m^{-3})
0
200
400
600
800
1000
0
2
4
6
8 x10^{10}
OH exposure (molecules cm^{-3} s)



(b) 2-heptenal
ΔLVP = 1.34
mfrag = 0.313
POx1 = 0.699
POx2 = 0.300
POx3 = 0.0004
POx4 = 0.0004
SOA concentration (μg m^{-3})
0
400
800
1200
1600
0
1
2
3
4
5
6
7 x10^{10}
OH exposure (molecules cm^{-3} s)

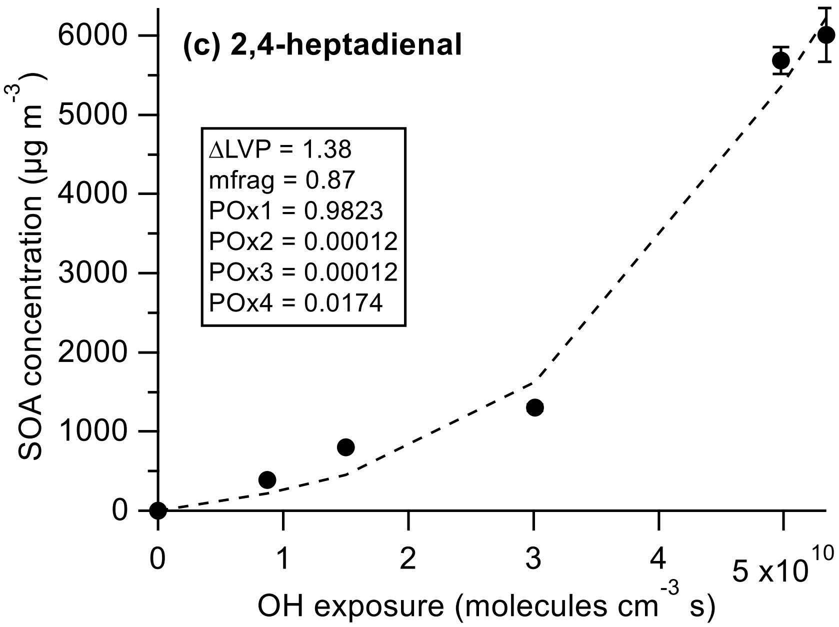


Figure S10. Estimation of SOM parameters by fitting SOA concentrations against OH exposure for heptanal (a), *trans*-2-heptenal (b), and *trans,trans*-2,4-heptadienal (c) photooxidation.

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