Reply for reviewer (1#)'s questions

We would like to thank the reviewer for the constructive comments. We really appreciate these comments as they will surely lead to improved manuscript. Below are the point by point the answers to the reviewer comments.

General comments:

In this paper, the authors use direct high resolution mass spectrometry and liquid chromatography with high resolution mass spectrometry to study the aging of a water soluble organic carbon mixture from biomass burning aerosol. They also study the photochemistry of phenol and guaiacol in the aqueous phase, for comparison to their ambient results. This work provides an interesting combination of experimental methods and laboratory analyses to probe the composition and evolution of water soluble organic compounds at the molecular-level. However, prior to acceptance, I recommend the following clarifications in the presentation and discussion of the data, as well as the following grammatical corrections.

Thanks for the positive comments on our submission. We would like to further improve our manuscript following the concrete suggestions.

Specific comments:

In general, I think the structure of this paper could be improved. It is not immediately clear whether the data discussed in section 3.1 (Mass spectral *characteristics of WSOC in biomass burning particulate) are from fresh particles. It seems to be discussing fresh particles, based on the titles of the other sections, but this should be made clearer in the text or in the section title.*

Here the data discussed in section 3.1 are combined to present field-collected aerosol samples, e.g., wheat straw burning aerosol (WSBA) samples before photolysis. We have re-phrased it in section 2.1 and 3.1, and deleted the "fresh" in the text.

Section 3.2 (Photochemical oxidation of phenols under laboratory conditions) shows interesting results, but these results could be linked more explicitly to the ambient data presented in the paper.

We have added a new section (i.e., section 3.3: Comparison of the photochemical products of phenolic compounds and the CHO composition in WSOC extracts from WSBA samples) to discuss the data comparison (including O/C, H/C and OSc) of WSBA samples and photochemical product samples (see lines 387-403).

There is a short discussion of a comparison between laboratory and ambient data (lines 379-387) and other discussions in the following section (3.3), but as a reader, I find this information hard to keep track of. Perhaps the paper could benefit from a specific section for lab ambient *intercomparisons.*

We added a new specific section for inter-comparisons between lab and ambient samples (see section 3.3 in revised manuscript).

Section 3.3.2 seems to be discussing photochemical stabilityâ˘A˘ consider labeling more clearly as such.

The section title is changed as "Presentation of photochemically stable organic species"

Section 3.3.3 seems to be discussing changes in composition as a result of different aging timesâ˘A˘ Tconsider labeling more clearly as such.

The section title is changed as section 3.4.3 "Comparison of time-profile mass spectra of CHO composition in WSOC extracts from WSBA samples" (line 541).

The authors use straw burning aerosol in this study. How representative is straw as a fuel in the particular region the authors are studying? How representative is this fuel more generally? This should be addressed.â˘A˘ Tthere are lots of different types of fuel that are burned and the choice of straw should be put into appropriate context.

According to Chinese government statistics (http://www.moa.gov.cn), there were 5.70 million acre and 14.2 million acre of wheat

crop in Hebei province and Henan province in summer of 2019, respectively. Our sampling region was in Wenxian, in Henan province and Daming in Hebei province, where the wheat are the main crops in summer. To facilitate subsequent planting and management, a large amount of wheat straw was burned directly by farmers in the fields during the harvest season in 2013. This was not controlled burning, but an illegal open burning with random, unordered and uncontrollable features. We have properly addressed this question in section 2.1 (line 131-136).

This study uses negative mode ionization only: how might this skew the types of compounds/compound classes identified? This should be discussed. Negative mode will ionize compounds that can be readily deprotonated, but what about compounds that are not easily deprotonated and may show up preferentially in positive mode ionization (e.g. compounds like amines)? How representative are the data in encapsulating mixture-wide characteristics if positive mode is not used?

Using negative mode ionization can identify those compounds that can be readily deprotonated, and our focus was on those compounds with common formulas including C, H, O, N and S atoms (see section 2.5). It is true that a large amounts of not easily deprotonated compounds will show up preferentially in positive mode ionization (e.g.

compounds like aldehydes, ketones, esters, amines), and they may also have different ionized forms (e.g., $[M+H]^+$, $[M+Na]^+$, $[M+K]^+$…) that makes them hard to be identified. Furthermore, compounds containing polyfunctional groups cannot be also easily identified in negative mode, but rather in positive mode. We have mentioned the limitation in the text (line 279-283).

The ionization method restricts the ability in encapsulating wide-mixture- characteristics of aerosols, because polar WSOC occupy only a fraction of aerosol composition, and uncertainties in those compounds (e.g., high molecular-weight matters) are out of detected range. Here we tentatively examine the molecular characteristics under our method limitation and cannot estimate the representation of identified compounds, especially lacking the data from positive mode detection.

Line 211: CHS compounds do not ionize well with electrospray ionization, which may explain why they are not detected with much prevalence here. This should be discussed.14

We have mentioned it at line 276-278.

Are the molecular formulas with the lowest ppm mass difference selected here? Are there any other QC/QA methods you use, like ensuring H/C ratios are reasonable or checking for non-integer DBEs in neutral *formulas?*

Here the assignment of molecular formulas were based on lowest ppm mass difference between the measured and the theoretical ion (see section 2.5). The formula assignment was based on the measured mass, before determination, the Orbitrap analyzer was externally calibrated for mass accuracy using Thermo Scientific Pierce LTQ Velos ESI calibration solution (composed of m/z 265 sodium dodecyl sulfate, m/z 514 iodine sulfonic acid sodium and m/z 1079-1979 Ultramark polymers). We excluded those neutral formulas with non-integer DBEs as they were not common compounds that might be ionized with unreasonable H/C or included other elements.

In general, I find the methods a bit challenging to follow. There are lots of interesting sampling and analytical methods used here and the differences between them for accurately interpreting the results are important. Perhaps the authors could include a summary table or flow chart of the different types of sample collection methods and data analysis methods for field and ambient samples.

We have supplemented the sampling and analytical methods in section 2.1:

Line 131-149: Field campaign. To facilitate subsequent planting and management, a large amount of wheat straw is burned directly in the

fields during the harvest season. The sampling fields were located in the north China plain, surrounded by wheat farm and far away from downtown. The smoke from open burning of wheat straw which enveloped the farm region was the main source of sampling particulate matter. (These fire events were not controlled burning process, but an illegal open burning with random, unordered and uncontrollable features. Because these events randomly occurred, the planning prior to events was not possible; hence, the samples reflect the real field situation.)

Line 152-157: Extract preparation.

Line 167-169 and 174-179: Photolytic experiment.

Line 228-231: Are these percentages by occurrence? By abundance? If weighted by abundance, how is ionization efficiency accounted for?

By occurrence, meaning assigned molecular formula species (see line 259: In the amount of assigned formulas…), no ion intensity.

In Figures 3 and 4, what do the boxes around the data mean? This should be mentioned in the figure caption.

We labeled the two areas with A and B, and added a sentence “Areas A and B are tentatively attributed to aliphatic and aromatic molecules, respectively.”, see Figure 2 (in section 3.1) that replaces previous

Figure 3.

The previous Figure 4 now is replaced by Figure 3 (in section 3.1), and the dashed area in previous Figure 4 was deleted.

In Figures 2-4, the authors show 1-2 representative samples. How do we know that these samples are truly representative? The authors should consider showing the rest of the data in the SI and highlighting how similar the data are, or perhaps should consider finding a way to show averages across all samples in main text figures.

We have re-arranged these figures. In the revised version, Figure 2 includes a reconstructed mass spectrum and VK diagram of one representative sample, and other three samples are shown in Figure S2.

The previous Figure 3 is deleted, instead of OSc including three samples (including laboratory samples). Other three field samples are shown in Figure S5.

Line 325-327: I agree that there is an increase in the average O/C of product compounds relative to their precursors, but this is to be expected. There are many products formed during phenol and guaiacol photochemistry, some of which probably have higher O/C than the precursor compounds and some of which may fragment and have lower O/C. While the average O/C of the product mixture increases, it should be acknowledged that this is an average and that many different product *compounds are formed.*

Reviewer is right. Here we use "average O/C" instead of "O/C" (see line 358).

Line 343-345: What were the starting pH values and are these changes statistically significant?

As described in section 2.3 (lines 189), the starting pH values were adjusted to 5, and these changes are statistically significant with T-test ($p<0.05$). We have rephrased the sentence (see line 373).

Figures 5 and 6 don't show drastically different information, they can probably be combined.

We have combined these two figures, see Figure 4 (in section 3.4) in the revised manuscript.

Lines 511-526: It should be clearer which time points you are referring to for these O/C and H/C ratios (t=0? T=4? T=12?).

T=0, as the irradiation time extended from 0 to 12h (see at line 556).

Line 547: Past work that has discussed the relatively short lifetime of CHON compounds like nitrophenols (order of hours, depending on conditions). In this manuscript, the authors mention that CHON compounds tended to exhibit good stability. What do you estimate the atmospheric lifetime of CHON compounds observed in your analysis to be *and how does this compare to past literature?*

It is true that the lifetime of compounds depend on reaction conditions. Here some CHON compounds seemed to exhibit good stability, which might be caused by a limited reaction due to the low mass content or light-shielding effect caused by other light-absorbing matter. We have mentioned it at line 523-525.

Technical corrections:

The quality of the writing in this manuscript should be improved upon prior to publication. These are some suggested edits, but in general the authors should carefully review the language in their manuscript.

We have edited carefully the language in the revised manuscript.

Abstract.

-Line 34: remove comma between "precursor" and "were"

Done (line 34).

-Line 35: remove comma after "both"

Done (line 39).

-Line 39: "of" instead of "on"

Done (line 43).

-*Line 41: “extracts” instead of “extract”*

Done (line 46).

-*Line 47: “has” instead of “have”*

Done (line 51).

Introduction

- *Line 53: space before “WSOC”*

Done (line 57).

- *Line 61: space before “Although”*

Done (line 65).

- *Line 66-67: for clarity, please define functional groups in words first before using abbreviated descriptions.*

Done (line 70).

- *Line 68: “carbonyls” instead of “carbonyl”.*

Done (line 72).

- *Line 71: could be clarified, what do you mean by “to affect aerosol evolution processes”?*

We use "to generate low-volatility species" to replace "to affect aerosol evolution processes" (see line 75).

- Line 82: consider removing the word "emerged", its meaning is unclear here.

The word "emerged" is deleted (line 84).

- Line 84: "environments" instead of "environment"

Done (line 86).

- Line 109-112: this sentence needs re-structuring for clarity: be more explicit and clear about why LC will help relieve ion suppression and how it will help identify ions (both are possible because of differences in LC retention time between compounds, but this should be made more obvious, for readers who do not frequently use LC techniques)

"HRMS coupled with LC might be another complementary powerful tool for relieving ion suppression due to its abilities to separate and analyze different kind of compounds with differences in LC retention time."(see line 112-114).

- Line 118: "were" instead of "was"

Done, we have rephrased the sentence (see line 124-128).

- Line 120: the word “reference” does not seem to be what you mean to say here, what about changing to “for comparison”?

We have rephrased the sentence (see line 124-128), and the words of “for comparison” were deleted.

Methods

- Line 158: What is this ratio in your extraction vs. in actual cloud water?

Cai et al (2018), and Li et al (2017) reported that the formation of cloud droplets is associated with ambient liquid water content (LWC) values and PM 2.5 concentrations. In their observation at Mountain Tai in China, the LWC and $PM_{2.5}$ concentrations during the cloud events ranged widely from 0.01 to 0.39 gm^{-3} and from 11.1 to 173.3 μg m^{-3}, respectively. The high LWC could facilitate the formation of larger cloud droplets, whereas higher PM2.5 levels resulted in higher concentrations of water-soluble ions. The calculated ratio of LWC/ PM 2.5 mass from Li et al (2017) was in a wide range from 1.4×10^2 to 1.6×10^4, which is comparable with the ratios of extracted water mass/ PM mass (M_{H2O}/ $PM_{2.5}$) ranging from 1.8×10^3 to 3.4×10^4 in the present study, indicating that the aqueous extract from this study is similar with the cloud water. Please see the lines 180-185 of the revised version.

- Line 160: "In the experimental section of phenol photochemistry" is confusing, perhaps re-phrase to "To study phenol photochemistry"

We have deleted it (line 187).

- Line 160: "solution" should be "solutions"

Done (line 187).

- Line 162: remove comma after "H2O2"

Done (line 188).

- Line 166-169: should be re-phrased, also I imagine you do not mean that you are looking for "biomarkers" but instead "tracer compounds"?

Done (line 194).

- Line 171: define "ESI" the first time you use it

Done (see at line 124).

- Line 176: Orbitrap should be capitalized consistently throughout the text

Done.

- Line 185: the plural of "spectrum" is "spectra"

Done (line 210).

- Line 208: please write out the full name of the “nitrogen rule”

Done (line 234).

Results and Discussion

- Line 215: title seems incomplete, missing the word “matter”

The title is replaced with: 3.1 Mass spectral characteristics of WSOC extracts from WSBA samples (line 241).

- Line 216: “2.5” should be written as a subscript here and throughout the text (i.e.PM2.5)

Done (see at line 141 and 242).

- Line 224: “spectrum” should be “spectra” if the authors meant it to be plural, also what is meant by “abstracted blank”? I assume this should say “subtracted”

Done (line 254).

- Line 234-235: this sentence is incomplete

We use “in spite of” instead of “although”, and remove the comma (see line 263).

- *Line 265: this sentence should be clarified: what is meant by "low content"? Low mass loading?*

Low mass loading (line 294).

- Line 290: remove "emerged"

Done (line 321).

- Line 320: remove comma after "4h".

Done (line 353).

- Line 328: add "after" before the word "photooxidation"

Done.

- *Line 321-331: should be clarified, if phenols and methoxyphenols are undergoing photochemical aging, how are they impacting POA?*

The paragraph is re-arranged (line 376-386), and the sentence is deleted.

- Line 351: remove comma after "mechanism".

Done (line 381).

- *Line 357: for clarity, consider using "chromatograms" instead of* *“diagrams”, and propagate through text.*

Done (line 404).

- *Line 378: remove “emerged”.*

Done (line 423).

- *Line 380: remove “stemmed”.*

Done (line 424).

- *Line 505: “parameter” should be plural.*

Done (line 548).

- *Line 520: should be re-phrased for clarity, what is meant by “increased tendency”, what exactly is “consistent with LC observations”?*

We have re-phrased it, see line 562-565.

Reply for reviewer (2#)’s questions

All authors would like to thank the reviewer for the constructive comments. These comments will surely improve our manuscript. Below are the point by point the answers to the reviewer comments.

This is an interesting manuscript that describes photo-oxidation of both representative "real" biomass burning organic aerosol (straw) as well as a simpler surrogate containing.

I find the paper hard to follow. This is in part because the written English, while passable, is imperfect. However, the main issue is that there is no clear story beyond "oxidation of WSOC increases the oxidation state" (which is almost guaranteed) and "even a model system from a single fuel is very complex".

We have realized the deficiency in writing and further improved the structure of this paper. The results and discussion include:

3.1 Mass spectral characteristics of WSOC extracts from WSBA samples.

3.2 Mass spectral characteristics of the products from photooxidation of phenolic compounds in the aqueous phase.

3.3 Comparison of the photochemical products of phenolic compounds and the CHO composition in WSOC extracts from WSBA

samples.

3.4 Photolysis of WSOC extracts from WSBA samples

As far as I can tell a manuscript with a more clearly articulated story would be appropriate for ACP, but I am also on the edge of the subject area, and so a reviewer with greater experience with high-resolution GC methods might be more appropriate.

Some general comments:

The "/" in "O/C" really means ratio, so "O/C ratio" is redundant. I suggest writing "the oxygen to carbon ratio (O/C)" once and then omitting "ratio" when subsequently using the abbreviation.

In this study, we write "the oxygen to carbon ratio (O/C)" once and then use O/C or O/C values to replace it .

There is room in the literature for complex analyses of complex systems and we can not always demand an incredibly simple story, but the paper could still benefit from a major re-write to pull the most important themes to the surface.

We have improved the structure of this paper, especially we added Section 3.3: Comparison of the photochemical products of phenolic compounds and the CHO composition in WSOC extracts from WSBA samples.

I do not believe it is appropriate to end the abstract with "accounting for the highly oxygenated nature". Perhaps "contributing to" is warranted but the implication of the current wording is that the contribution dominates, and that has not been demonstrated here.

Done. We use "contributing to" to replace "accounting for" at the end of the abstract.

Specific comments:

Line 252 "is prone for" is not quite right. "is most sensitive to molecules containing polar ..."?

We have rephrased the sentence (see line 279-283).

Line 287 "the all extract samples" ???? Either the authors actually mean "the all-extract samples" or they may mean "all of the extracted samples". Clarify.

Here we have deleted the word of "all" (see line 318).

Line 290 "the emerged O/C ratios" could be "the measured O/C values" ("values" is appropriate after O/C in my opinion).

Here we use "the measured O/C" to replace "the emerged O/C ratios" (see line 321).

Line 336 "as well as to increased" clashes with the subject "would result in" before, so "to" should be "in".

Done (see line 365).

Line 347 "pathway for the low-volatility" strike "the".

Deleted (see line 377).

Line 532 "experience the similar" again strike "the".

Deleted (see line 582).

Line 543 "aerosols have the potential" the subject is "fraction" so should be "has".

Done (see line 596).

Line 544 "partly account for" is better than the abstract but "contribute to" would be best in both places.

Here and the abstract we use "contribute to" (see line 52 and 597).

Molecular Composition and Photochemical Evolution of Water Soluble Organic Carbon (WSOC) Extracted from Field Biomass Burning Aerosols using High Resolution Mass Spectrometry

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ABSTRACT

Photochemistry plays an important role in the evolution of atmospheric water soluble organic carbon (WSOC), which dissolves into clouds, fogs and aerosol liquid water. In this study, we tentatively examined the molecular composition and evolution of a WSOC mixture extracted from field-collected wheat straw burning aerosol (WSBA) samples upon photolysis, using direct infusion electrospray ionization (ESI) coupled to high-resolution mass spectrometry (HRMS) and liquid chromatography (LC) coupled with HRMS. For comparison, two typical phenolic compounds (i.e., phenol and guaiacol) emitted from lignin pyrolysis in combination with hydrogen peroxide (H_2O_2) as a typical OH radical precursor were simultaneously exposed to simulated sunlight irradiation. Their photochemical products such as phenolic dimers (e.g., m/z 185.0608 for phenol dimer and m/z 245.0823 for guaiacol dimer) or their isomers, were also observed in field-collected WSBA samples, suggesting that the aqueous-phase reactions might contribute to the formation of emitted biomass burning aerosols. The aqueous photochemistry of both the phenols (photooxidation) and WSBA extracts (direct photolysis) could produce a series of highly oxygenated compounds which in turn increases the oxidation degree of organic composition and acidity of the bulk solution. In particular, the LC/ESI-HRMS technique revealed significant photochemical evolution of the WSOC composition in WSBA samples, e.g., the photodegradation of low oxygenated species and the formation of highly oxygenated products. We also tentatively compared the mass spectra of photolytic time-profile WSBA extracts with each other for a more comprehensive description of the photolytic evolution. The calculated average oxygen-to-carbon ratio (O/C) of oxygenated compounds in bulk extract increases from 0.38 ± 0.02 to 0.44 ± 0.02 (mean ± standard deviation) while the intensity (S/N)-weighted average O/C (O/C_w) increases from 0.45 ± 0.03 to 0.53 ± 0.06 as the time of irradiation extends from 0 to 12h. These findings indicate that the water soluble organic fraction of combustion-derived aerosols has the potential to form more oxidized organic matter, ~~accounting for~~contributing to the highly oxygenated nature of atmospheric organic aerosols.

1 INTRODUCTION

Water-soluble organic carbon (WSOC) comprises a significant fraction of atmospheric aerosols, accounting for 20–80% of total organic carbon (OC) (Krivacsy et al., 2001; Wozniak et al., 2008; Fu et al., 2015; Xie et al., 2016). WSOC is directly involved in the formation of cloud condensation nuclei (CCN) by modifying the aqueous chemistry and surface tension of cloud droplets (Graham et al., 2002; Nguyen et al., 2012; Zhao et al., 2013; McNeill 2015). Despite its significance, little is known about the chemical composition and sources of WSOC, with less than 10–20% of the organic mass being structurally identified (Cappiello et al., 2003; Fu et al., 2015). Biomass burning is a well-known emission source of WSOC (Anastasio et al., 1997; Fine et al., 2001; Graham et al., 2002; Mayol-Bracero et al., 2002; Gilardoni et al., 2016). Although the composition varies with fuel type and combustion conditions (Simoneit 2002; Smith et al., 2009), the WSOC mixture often covers a common range of polar and oxygenated aromatic compounds (Graham et al., 2002; Mayol-Bracero et al., 2002; Duarte et al., 2007; Chang and Thompson 2010; Yee et al., 2013; Gilardoni et al., 2016) with molecules incorporating different numbers of functional groups like hydroxyl, carboxyl, aldehyde, ketone, ester, amino and/or other nitrogen-containing groups (Graham et al., 2002). In particular, lignin pyrolysis often yields a large amount of aromatic alcohols, carbonyls, and acid compounds (Mayol-Bracero et al., 2002; Chang and Thompson 2010; Gilardoni et al., 2016). Once dissolved into cloud, fog, and even aerosol liquid water, these substances can undergo aqueous-phase reactions to generate low-volatility species under sunlight irradiation, which have the potential to form secondary organic aerosol (SOA) after water evaporation (Graham et al., 2002; Cappiello et al., 2003; Duarte et al., 2007; Sun et al., 2010; Yu et al., 2014).

Field and laboratory studies have demonstrated that aqueous photochemical processes contribute significantly to the aqueous SOA formation from biomass burning precursors and the evolution of smoke particles (Sun et al., 2010; Lee et al., 2011; Kitanovski et al., 2014; Yu et al., 2014; McNeill 2015; Gilardoni et al., 2016). Gilardoni et al. (2016) observed aqueous SOA formation in both fog water and wet aerosols, resulting in an enhancement in the oxidized OA, and following atmospheric aging the overall oxidation degree of aerosols has also increased. In laboratory studies, phenols and methoxyphenols (important biomass burning intermediates) are often used as SOA precursors to examine the photochemical evolution in aqueous environments and aerosol-forming potential under relevant atmospheric conditions (Chang and Thompson 2010; Sun et al., 2010; Smith et al., 2014; Yu et al., 2014; Vione et al., 2019). The corresponding photochemical products formed through hydroxylation, oligomerization, and fragmentation typically cover a series of low-volatility and highly oxygenated species. For instance, the methoxyphenol-derived SOA are proposed as a proxy for atmospheric humic-like substances (HULIS) (Ofner et al., 2011; Yee et al., 2013). Other compounds emitted from lignin pyrolysis, e.g., aromatic alcohol, carbonyl, and carboxylic species retaining the phenyl ring have also been found to produce colored products via aqueous photooxidation, which may become a part of HULIS (Chang and Thompson 2010; Huang et al., 2018). In addition, photochemical processing of common water-soluble aliphatic compounds such as aldehydes (Lim and Turpin 2015), polyols (Daumit et al., 2014), and organic acids (Griffith et al., 2013) in aqueous solution can also lead to the formation of oligomers, highly oxygenated and multifunctional organic matter (McNeill 2015).

In recent years, high resolution mass spectrometry (HRMS) has been commonly applied to study the organic molecular composition in cloudwater (Zhao et al., 2013; Boone et al., 2015), fogwater (Cappiello et al., 2003), rainwater (Altieri et al., 2009a; Altieri et al., 2009b), laboratory-generated SOA (Bateman et al., 2011; Romonosky et al., 2015; Lavi et al., 2017), and field-collected aerosol samples (Laskin et al., 2009; Lin et al., 2012a; Lin et al., 2012b; Kourtchev et al., 2013; Tong et al., 2016; Wang et al., 2017). It has also been used in time-profile observations of the photochemical evolution of aqueous extracts from laboratory-generated SOAs (Bateman et al., 2011; Romonosky et al., 2015). However, direct infusion MS methods are prone to ion suppression caused by other organic species, inorganic salts, and adduct formation (Kourtchev et al., 2013). Therefore, HRMS coupled with liquid chromatography (LC) might be another complementary powerful tool for relieving ion suppression due to its abilities to separate and analyze different kind of compounds with differences in LC retention time (Kourtchev et al., 2013; Wang et al., 2016). It could also provide more information enabling the identification of possible isomers from the ions with same mass-to-charge ratio (m/z).

To our knowledge, the aqueous photochemical evolution of WSOC extracted from real ambient aerosols has not been studied in detail at the molecular level. Our previous study has revealed that the ultraviolet-visible (UV-VIS) absorption spectra of aqueous extracts from field biomass burning aerosols were modified under simulated sunlight illumination (Cai et al., 2018). Based on the previously studied field-collected samples, the present study is focused on a further analysis to investigate the molecular characteristics of water-soluble organic molecules by the photochemical evolution using electrospray ionization (ESI)-HRMS and LC/ESI-HRMS performed in negative ionization mode. For comparison, we also evaluated the photochemistry of phenol and guaiacol (representing the basic structures of phenols emitted from lignin pyrolysis) under laboratory conditions, and tentatively traced some of their photochemical products (e.g. dimers) in field-collected samples under study~~used as a reference~~.

2 EXPERIMENTAL SECTION

2.1 Particulate sample collection and preparation of aqueous extracts

The wheat straw burning aerosol (WSBA) samples were collected during the summer harvest season of 2013, at rural fields in the plain of north China where the wheat was the main agricultural crop (Cai et al., 2018). To facilitate subsequent planting and management, a large amount of fresh wheat straw was directly burned in the field during the harvest season, and the water emitted from burning plant body could provide a suitable environment for aqueous photochemistry of dissolved compounds. The selected WSBA samples used for HRMS analysis were collected from two sampling sites, located at rural fields in Wenxian in Henan Province (noted: HNWX) and Daming in Hebei Province (HBDM). As described in Cai et al. (2018), the selected sampling sites were mainly affected by heavy smog from wheat straw burning (Figure 1). The emitted fine particulate matter with aerodynamic diameter $\leqslant 2.5$μm ($PM_{2.5}$) was collected at a flow rate of 5 L min^{-1} by a portable particulate sampler (MiniVol TAS, AirMetrics, USA), with quartz fiber filters (47mmin diameter, QMA, Whatman, UK) baked at 600°C for 6 hours before sampling. The sampling flow rate was calibrated with a standard flow meter (Bios Defender 520) and the sampling time of each filter was restricted to 30-60 minutes depending on the ambient biomass burning aerosol concentration and expected filter loading (Cai et al., 2018). After collection, the filter samples were stored in dark and transported to the laboratory, and then stored at -20°C under a light-proof condition.

The preparation of WSOC extracts and measurements for carbon content including organic carbon (OC), elemental carbon (EC) and WSOC were described in detail in Cai et al. (2018). Briefly, a part of each quartz fiber filters (1.6-3.2cm^2) was placed into a brown vial and extracted with ultra-pure water (Milli-Q, Milipore) for two times; at each time 5 ml ultra-pure water with a 30min ultrasonic agitation was applied. The two-time extracts were combined and filtered through a PTFE syringe filter (0.2μm pore size, Thermo Scientific), followed by a pH measurement with a pH meter (Mettler Toledo SevenEasyTM S20) that has been regularly calibrated at pH 4.00 and 6.86. Prior to analysis the extracts were stored at -20°C in the dark. To reduce the WSOC mass loss, the desalting treatment (e.g., solid phase extraction (SPE)) was not performed on these samples.

Figure 1. One field site at Daming, Hebei province, China, for sampling the aerosols affected by biomass burning.

2.2 Direct photolysis of WSOC extracts

A 12-hour direct photolysis of WSOC extracts obtained from WSBA samples was performed in a photo-reactor (BL-GHX-V, Bilon Instruments Co. Ltd., China, see Figure S1) that was equipped with a solar simulator (Xe lamp, 1000W) placed into a double-deck quartz condenser (Cai et al., 2018). A cooling water (18°C) was circulating in the outer tube of the condenser to avoid heating of the samples. In the wavelength range of 310-400 nm relevant to the boundary layer of the atmosphere, the actinic flux of the lamp is about 5 times stronger than the solar actinic flux, meaning that the spectral evolution via the 12-hour simulated solar irradiation might be equal with the effect caused by actual sunlight irradiation with a duration of at least 60 hours (Cai et al., 2018). Air-tight quartz tubes (1.5cm in diameter, 3ml solution per tube) loading extracts were equidistantly arranged around the lamp. Each extract was distributed into three tubes that corresponded to three different irradiation times, i.e. 0, 4, and 12 h, with no oxidants added externally throughout the whole photolytic process. At each irradiation time point (e.g., 0 and 4 h), the related tubes were wrapped with aluminum foil, and placed at the initial location until the end of 12-h photolysis (Cai et al., 2018).

As described in Cai et al. (2018), the water extraction resulted in a dilution of the collected organic compounds, however, the ratio of the water mass to $PM_{2.5}$ mass for extract samples (ranging from 1.8×10^3 to 3.4×10^4) was compatible with the ratio of water mass to WSOC content in cloud water (in a wide range from 1.4×10^2 to 1.6×10^4) (Li et al., 2017), indicating that the present aqueous extracts are relevant to the atmospheric cloud water condition.

2.3 Photooxidation of phenolic compounds under laboratory conditions

Initial solutions of 0.1 mM phenol (C_6H_6O) and 0.1 mM guaiacol ($C_7H_8O_2$) in combination with an OH radical precursor (0.1 mM H_2O_2) were prepared in ultra-pure water (Milli-Q, Milipore). The pH of the solution was adjusted to 5 with 0.1 M sulfuric acid (H_2SO_4), which is usually relevant to the acidity in fog and cloud waters (Collet et al., 1998; Fahey et al., 2005). The prepared solution and reference blank were irradiated by simulated sunlight irradiation with a duration of 4 hours. Hereby, we mainly focus on acquiring the chemical characteristics of aqueous products of phenols, and tentatively identify some tracer compounds (e.g., phenolic dimers) whether they exist in present biomass burning particulate samples.

2.4 Sample analysis

The direct infusion MS analysis was conducted using a Thermo Scientific Orbitrap Fusion Tribrid mass spectrometer equipped with quadrupole, orbitrap, and linear ion trap mass analyzers, with a heated ESI source. To assist in ionization and desolvation, the sample was diluted to a 1:1 mixture of acetonitrile and sample by volume. The full scan mass spectra were acquired in negative ionization mode, with a resolution of 120 000 at *m/z* 200 for the Orbitrap analyzer and a mass scan range of m/z 50-750. Before determination, the Orbitrap analyzer was externally calibrated for mass accuracy using Thermo Scientific Pierce LTQ Velos ESI calibration solution. The direct infusion parameters were as follows: sample flow rate 5μl min^{-1}; capillary temperature 300°C; S-lens RF 65%; spray voltage -3.5kV; sheath gas, auxiliary gas, and sweep gas flows were 10, 3, and 0 arbitrary units, respectively. Data collecting was performed when the intensity of the total ion current (TIC) maintained constant with a relative standard deviation (RSD) under 5%. At least 100 data points (mass spectral scans) were collected for each test sample, and the each exported mass spectrum for analysis was derived from the average result of 100 spectrums.

The LC/ESI-HRMS analysis operated in negative ionization mode was performed using a U3000 system coupled with a T3 Atlantis C18 column (3μm; 2.1×150mm; Waters, Milford, USA) and an Orbitrap Fusion MS. A 10 μL sample was injected, with a flow rate of 0.2 ml min^{-1} for the mobile phase, which consisted of H_2O (A) and acetonitrile (B). The gradient applied was 0-5 min 3% B; 5-20 min from 3 to 95% (linear), and kept for 25 min at 95%; and 45-50 min from 95 to 3%, and held for 10 min at 3% (total run time 60 min).

2.5 Data processing

Mass spectral peaks with three times larger than the signal to noise ratio (S/N) were extracted from the raw files. Peaks in both sample and blank spectra were retained if their intensity in the former was five times larger than in the latter. A common molecular assignment based on the accurate mass was performed using Xcalibur software (V3.0 Thermo Scientific) with the following constraints: $^{12}C \leq 50$, $^{13}C \leq 1$, $^{1}H \leq 100$, $^{16}O \leq 50$, $^{14}N \leq 4$, $^{32}S \leq 1$, and $^{34}S \leq 1$. All mathematically possible elemental formulas, with a mass tolerance of ±3ppm were calculated. Elemental formulas containing ^{13}C or ^{34}S were checked for the presence of ^{12}C or ^{32}S counterparts, respectively. If they were not matched with the corresponding monoisotopic formulas, then the assignment with next larger mass error was considered. Isotopic and unassigned peaks were excluded from further analysis.

Ions were also characterized by the number of rings plus double bonds (i.e., double bond equivalents (DBE)), which were calculated as: $DBE = c - h/2 + n/2 + 1$ for an elemental composition of $C_cH_hO_oN_nS_s$. The assigned formula was additionally checked with the nitrogen rule. For ambient samples, based on the presence of various elements in a molecule, the identified elemental formulas were classified into several main compound classes: CHO (i.e., molecules containing only C, H, and O atoms), CHOS, CHON, and CHONS, and others including CHN and CHS. In the present study, because the detected water-soluble ions almost were below m/z 400, we focused our molecular analysis on m/z 50-400.

3 RESULTS AND DISCUSSION

3.1 Mass spectral characteristics of WSOC extracts from WSBA samples

The preliminary analysis showed that the $PM_{2.5}$ concentration in ambient air near to the burning sites ranged from 6.46 to 28.03 mg m^{-3} (Table S1). OC was the major component of the collected $PM_{2.5}$ with a proportion of 50.9±7.6% (mean±standard deviation), whereas EC represented a negligible fraction (average 1.3 ± 0.4%). Meanwhile, WSOC accounted for 35.5±7.5% of OC in the tested samples.

Although this batch of aerosol samples were collected from different sites, their water-extracted solutions showed similar light-absorbing characteristics in UV-VIS absorption spectra (Cai et al., 2018). Here, four extract samples (HNWX-1, HNWX-2, HBDM-1 and HBDM-2) (Table S1) were chosen for further analysis using high resolution mass spectrometry. These samples also exhibited similar patterns in mass distribution of water-soluble molecular species that mainly range from 50 to 400 Da, which indicated a similar burning source for these samples. A reconstructed mass spectrum (subtracted blank) for one representative sample of HNWX-1 is shown in Figure 2a (others are shown in Figure S2). In mass range 50-400 Da, there were 827±44 molecular formulas identified throughout the all samples, and most of the formulas (above 75%) were overlapped between these analyzed samples. The classification features of assigned compounds for analyzed extracts are shown in Table S2. In the amount of assigned formulas, CHO composition was the most abundant group, accounting for 59.2±2.2% of the total assignments, followed by CHON (35.0±2.2%). These results are consistent with previous observations of laboratory-generated biomass burning aerosol (Smith et al., 2009) and field particulate samples influenced by biomass combustion (Kourtchev et al., 2016) in spite of the differences of biomass varieties, extracted solvents, and HRMS techniques between present and previous studies.

On the other hand, CHOS and CHONS compounds contributed with less than 5% to the total assignment. A number of studies have shown the wide presence of organosulfates and nitrooxy-organosulfates in urban (Lin et al., 2012b; Wang et al., 2016), rural (Lin et al., 2012a), and forest aerosols (Kourtchev et al., 2013), and even in cloudwater (Boone et al., 2015); however, most of these compounds were not observed in our negative mass spectra. This could be accounted for by the low extent of aerosol evolution, due to the limited oxidation conditions available for the formation of organosulfates and nitrooxy-organosulfates in fresh smoke aerosols. For example, laboratory studies have observed the significant formation of organosulfates via photooxidation in the presence of acidic sulfate aerosol (with significant level of SO_2 concentration) (Surratt et al., 2007; Surratt et al., 2008). All detected ion species with enabled formula assignments in present samples are listed in Table S3. The electrospray ionization did not ionize well CHN and CHS compounds; hence, they were not the most prevalent compounds in present samples.

It should be also noted that the negative ionization mode prefers to detect those molecules containing polar functional groups (e.g., -OH and –COOH) that could be readily deprotonated. There were number of compounds that were not easily deprotonated and might show up preferentially in positive ionization mode (e.g. amines). Furthermore, the formula numbers detected in the HRMS potentially contain multiple structural isomers; therefore, the actual number of water-soluble organic species is expected to be underestimated. The additional LC/ESI-HRMS analysis operated in negative mode confirmed a substantial number of ion masses (e.g., assigned CHO and CHON compounds) containing more than one structural isomer, which could be observed at different retention times (RTs) in chromatograms. Two representative groups of extracted chromatograms for CHO ($[C_7H_5O_n]^-$, (n=2~4)) and CHON ($[C_7H_5O_nN]^-$, (n=1~3)) compounds are shown in Figure S3 and S4, respectively, where increasing the O or N atom number in a molecule might lead to more isomer peaks. However, it should be noted that these LC-separated peaks might also include other unidentified compounds that were outside of the elemental assignment considered in this study. Additionally, low mass loading and potential decomposition under the ionization can also limit the detection of some high molecular weight species.

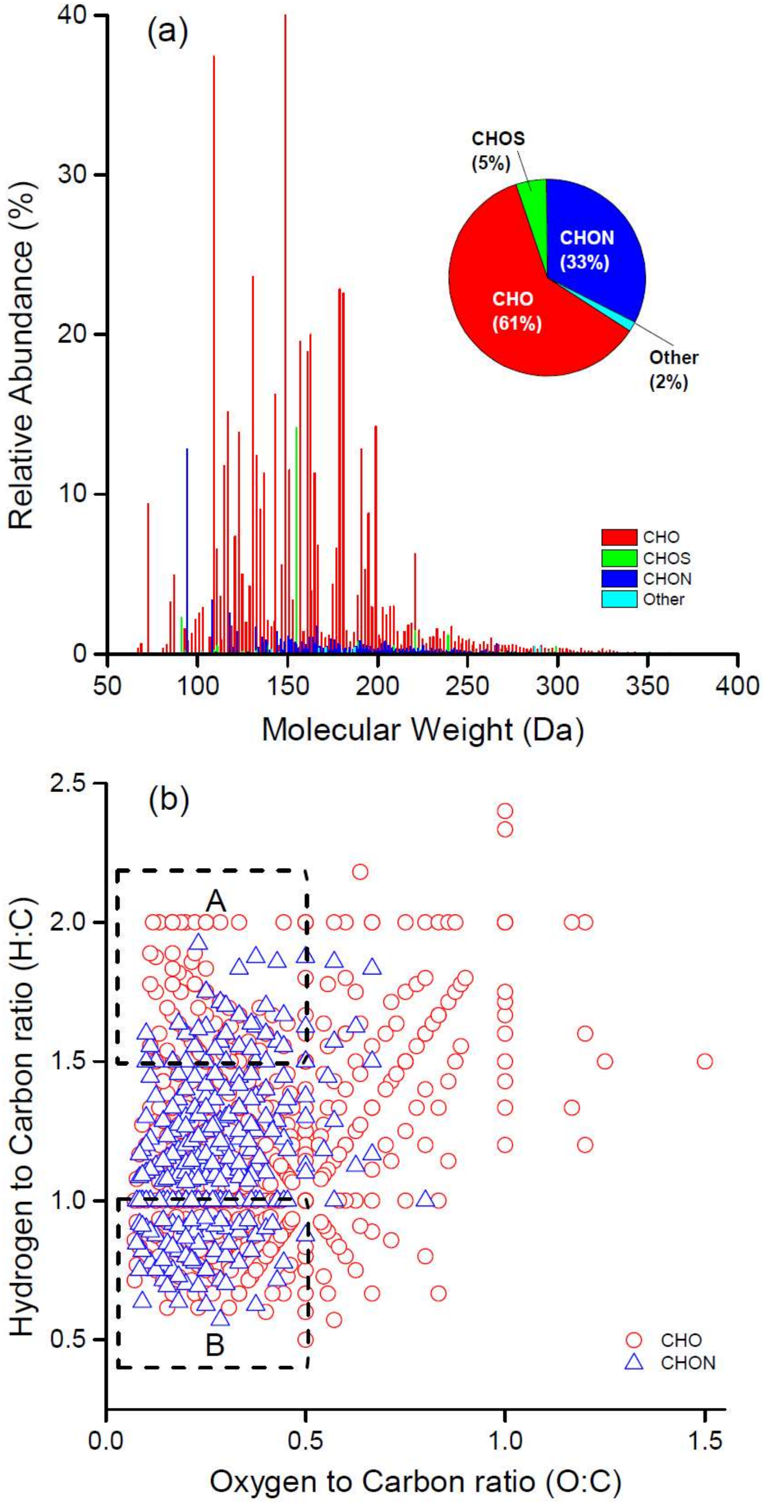

Figure 2. (a) Reconstructed mass spectra for detected ions with assigned formulas and (b) Van Krevelen diagrams for CHO and CHON species in extract of HNWX-1 sample. The inset pie charts in (a) show the number fraction of each class in the total assigned compounds. Areas A and B in (b) are tentatively attributed to aliphatic and aromatic species, respectively.

The interpretation of the complex organic mass spectra generated by high resolution mass spectrometry can be simplified by plotting the hydrogen to carbon ratio (H/C) against the oxygen to carbon ratio (O/C) for individual assigned atomic formulas in form of the Van Krevelen (VK) diagram (e.g. Lin et al., 2012a; Kourtchev et al., 2013). Figure 2b indicates a representative VK diagram of CHO and CHON compounds derived from HNWX-1 sample. It can be clearly seen from Figure 2b that the majority of CHO and CHON molecules are located at the region of O/C⩽1.0 and H/C⩽2.0. In VK diagram, molecules with H/C⩽1.0 and O/C⩽0.5 are typical for aromatic species, while molecules with H/C⩾1.5 and O/C⩽0.5 would be associated with typical aliphatic compounds (Mazzoleni et al., 2012; Kourtchev et al., 2014). The average double bond equivalent (DBE) showed relative high values with 5.5 for CHO compounds and 6.1 for CHON compounds (Table S2), suggesting that oxidized aromatic compounds were abundant in the present sample, and their presence could partially account for the strong light-absorbing feature in the near-UV region as observed in our previous study (Cai et al., 2018).

The average H/C and O/C values throughout the extract samples were in the ranges of 1.26-1.31 and 0.34-0.42 for CHO compounds, 1.19-1.23 and 0.28-0.29 for CHON compounds (shown in Table S2), respectively. Although the ESI analysis were performed in the negative ionization mode, the measured O/C exhibit rather low values, which fall in the range of O/C ratios typical for biomass burning organic aerosol derived from positive ionization mode (Aiken et al., 2008; Kourtchev et al., 2016). Due to fresh emission and smaller aging effect, the present O/C were obviously lower than the O/C of long-range transport biomass burning aerosols (Zhang et al., 2018).

Carbon oxidation state (OS_c) was observed to increase with oxidation for atmospheric organic aerosol and link strongly to aerosol volatility (Kroll et al., 2011). OS_c for each molecular formula can be calculated using the following equation:

$$OS_c = -\sum_i OS_i \frac{n_i}{n_c}$$

where OS_i is the oxidation state associated with non-carbon element i and n_i/n_C is the molar ratio of element i to carbon within the molecule (Kroll et al., 2011; Kourtchev et al., 2013).

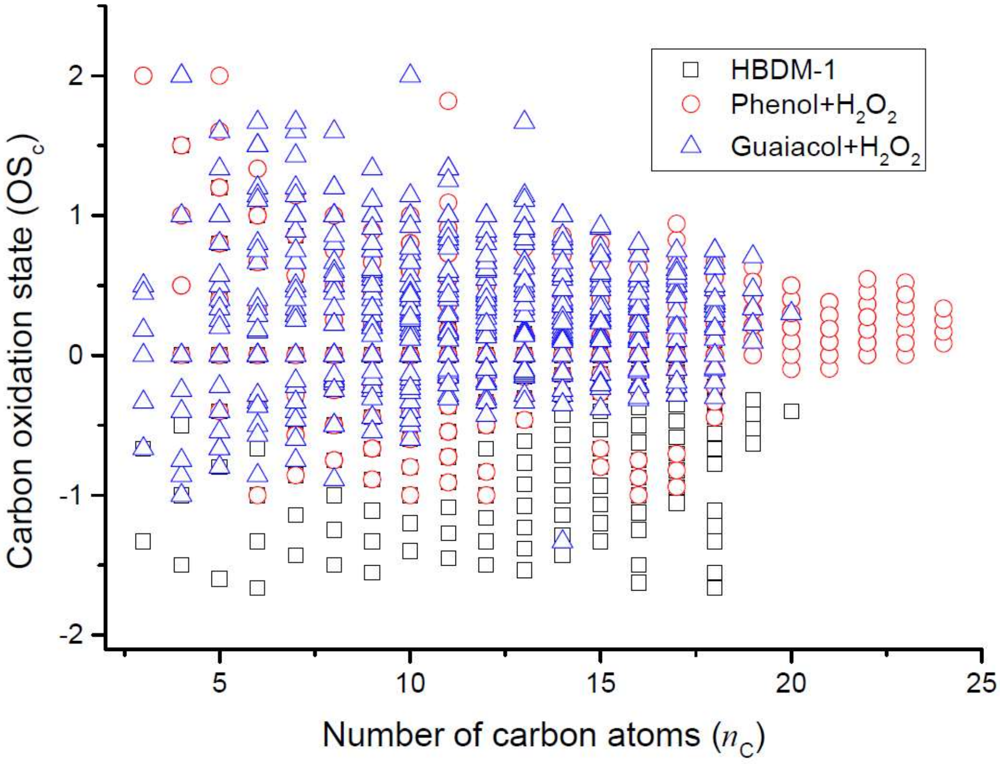

Figure 3. The distribution of carbon oxidation state (OS_C) for CHO molecules in HBDM-1 and laboratory samples.

Considering that nitrogen and sulfur atoms can present multiple oxidation states, the OSc was calculated and analyzed only for CHO compounds in this study. A similar pattern of OSc values versus the number of carbon atoms (n_C) was observed for CHO compounds detected in present WSBA samples (Figure 3 and Figure S5). From Figure 3 and Figure S5, it can be seen that OSc of each sample ranges mainly from -1.5 to +1 with average ranging from -0.6 to -0.4. Consistent with previous studies (Kroll et al., 2011; Kourtchev et al., 2016), the majority of molecules with $OS_C<0$ (low oxidized organics) and carbon atoms lower than 20 are suggested to be associated with the primary organic aerosols emitted from biomass burning. A minor fraction of molecular formulas with $OS_C\geq0$ values might be associated with semivolatile and low-volatility oxidized organic aerosols (Kroll et al., 2011). Figure 3 also shows the plot of OS_c versus n_C for products obtained from photooxidation of phenol and guaiacol, respectively, and their comparison with WSBA samples will be discussed in section 3.3.

3.2 Mass spectral characteristics of the products from photooxidation of phenolic compounds in the aqueous phase

Phenol and guaiacol were chosen as two representative model compounds derived from biomass combustion. Two high resolution mass spectra of aqueous phenol and guaiacol exposed to OH radicals for 4h are shown in Figure S6, where 435 $C_xH_yO_z$ molecular formulas (m/z 90-500) were assigned for product ions of phenol (with C_3-C_{24}) and 624 $C_xH_yO_z$ formulas (m/z 90-600) were assigned for product ions of guaiacol (with C_3-C_{27}). The average H/C and O/C values were 0.79±0.28 and 0.52±0.23 for phenol, and 0.88 ± 0.24 and 0.59 ± 0.24 for guaiacol, respectively. Clearly, the photochemical processing induced by OH oxidation resulted in an increase in average O/C of product molecules relative to their precursors (O/C=0.17 for phenol and O/C=

0.29 for guaiacol).

The formation mechanisms of series of oxygenated products, e.g., phenolic oligomers, hydroxylated phenolic species, ring-opening and highly oxygenated compounds, are proposed in the literature (e.g. Sun et al., 2010; Chang and Thompson, 2010; Yu et al., 2014; Huang et al., 2018). The OH-initiated reactions would result in enhanced hydroxylation of the aromatic ring as well as in increased yields of carboxylic acids and toxic dicarbonyl compounds (Sun et al., 2010; Yu et al., 2014; Prasse et al., 2018). For example, some highly oxygenated C_2-C_5 aliphatic compounds (e.g., $C_2H_2O_4$, $C_3H_4O_4$, $C_4H_6O_4$, and $C_5H_6O_5$) corresponding to carboxylic acids (Yu et al., 2014) were clearly observed in the mass spectra of present photochemical products. The occurrence of these oxygenated products not only directly increased the degree of oxygenation in the bulk solution composition, but also contributed to the variation of solution acidity. After the 4-hours photochemical process, the pH values of the irradiated solution were significantly lower than the pH values of the solution prior to irradiation (t-test, $p<0.05$), and the calculated acidities ($[H^+]$) of the bulk solution increased by $(2.96 \pm 0.15) \times 10^{-5}$ M and $(4.26 \pm 0.16) \times 10^{-5}$ M for phenol and guaiacol, respectively.

The oligomerization induced by photochemical transformation of phenolic substances is an important formation pathway for ~~the~~ low-volatility, light-absorbing compounds (Smith et al., 2016). Here, phenolic dimmers (i.e., $C_{12}H_{10}O_2$ for phenol dimer and $C_{14}H_{14}O_4$ for guaiacol dimer) and higher oligomers (e.g., $C_{18}H_{14}O_3$ and $C_{24}H_{18}O_4$ for phenol trimer and tetramer, $C_{21}H_{20}O_6$ for guaiacol trimer), as well as their hydroxylated species were observed. The formation mechanism can be ascribed to C-

O or C-C coupling of phenoxy radicals that were formed via H-abstraction of the phenols or OH addition to the aromatic ring (Net et al., 2009, Sun et al, 2010). The reaction at the para position or para-para coupling was more likely to occur due to a higher probability of free electron to occur in this position (Lavi et al, 2017) or a weaker steric hindrance in the para position.

3.3 Comparison of the photochemical products of phenolic compounds and the CHO composition in WSOC extracts from WSBA samples

Compared to the CHO compounds detected in WSOC extracts, the photochemical products of the two phenols under study showed a higher O/C and a lower H/C values. The average OS_C of photochemical products from phenol (OS_C=-0.7) and guaiacol (OS_C=-0.6) after a 4-hour photooxidation raised to +0.2 and +0.3, respectively, showing distinctly a higher degree of oxidation than the present WSBA samples. In Figure 3, more species with OS_C<0 (especially OS_C<-0.5) are presented in the field sample (HBDM-1), while the species with OS_C⩾0 are prevalent in photochemical products of phenol and guaiacol. The single-precursor systems in laboratory did not completely reflect the CHO composition features in water-soluble extracts from real straw-burning samples that contained a myriad of precursors and unknown substances from atmospheric background, soil and other sources. Considering that a large number of phenols and methoxyphenols exist in the straw-burning smokes and their potential to undergo photochemical aging, the nature of emitted primary organic aerosols is reasonably more complicated than the nature of simulated products derived from single-precursor systems.

The extracted LC chromatograms of m/z 185.0608 and 245.0823 are shown in Figure 4, respectively, where both ions involve dimers of phenol and guaiacol with several structures, and/or other isomers. The presence of guaiacol dimer and syringol dimer was previously observed in aerosol samples largely affected by wood combustion. Based on the Aerosol Mass Spectrometer (AMS) analysis, these two dimers were suggested as markers of biomass burning aerosols (Sun et al., 2010; Yu et al., 2014). In the composition of present biomass burning aerosols, the phenolic dimers (m/z 185.0608 and 245.0823) were also observed in present mass spectra, but the extracted LC chromatograms shown in Figure 4 indicate that these ions contain multiple RT peaks. The same peaks with RT18.3 and 19.2 min which are assumed to be the phenol dimers were observed during the photochemical transformation of phenol (Figure 4a) and in the WSBA samples. Meanwhile, the present particle extracts may also involve guaiacol dimer, since its m/z 245.0823 has two LC peaks emerged at RT 17.7 and 19.5 min (Figure 4b) same as the peaks identified during the photochemical transformation of guaiacol. Considering a large amount of water emitted during the process of straw combustion, the occurrence of phenolic dimers might indicate that the aqueous phase reactions played an important role in the formation and evolution of emitted aerosol organic composition.

Typical hydroxylated species such as, e.g., $C_2H_2O_4$, $C_6H_6O_2$, $C_7H_6O_3$, $C_7H_8O_3$, were also found in the samples ~~emerged~~ from photooxidation of both phenols and the WSBA samples. The comparison of the photochemical products ~~stemmed~~ from phenols and the WSBA samples revealed their significant difference, pointing to the importance of studying real aerosol samples against the laboratory model compounds. However, evaluating the model compounds as proxy of real aerosol samples is always helpful as a reference. To this end, it is worth noting that potentially other phenols and methoxyphenols (e.g., acetosyringone, vanillin) that dissolve into cloud, fog droplets or aerosol liquid water can be photochemically transformed and contribute to the SOA formation (Vione et al., 2019, Zhou et al., 2019).

3.4 Photolysis of WSOC extracts from WSBA samples

Although the direct photolysis was performed on present WSOC extracts from WSBA samples in presence of simulated sunlight irradiation without adding any oxidants, the photooxidation process still occurred since the particle extracts were very likely to include various oxidants, e.g., singlet molecular oxygen ($^1O_2$), peroxides, hydroxyl radical (OH) or excited triplet state of organics produced under light excitation (Anastasio et al., 1997; Vione et al., 2006; Net et al., 2009; Net et al., 2010a; Bateman et al., 2011; Rossignol et al., 2014; Smith et al., 2014; Gómez Alvarez et al., 2012). In particular, the excited triplet state of aromatic carbonyls (e.g., 3, 4-dimethoxybenzaldehyde) (Net et al., 2010b) was found to be more efficient than OH radical to oxidize phenols and produce hydroxylated species (Smith et al., 2014., Yu et al., 2014). This photosensitized reaction is likely to play an important role in the WSOC evolution, due to high quantities of aromatic carbonyls present in the extracts of biomass burning aerosols.

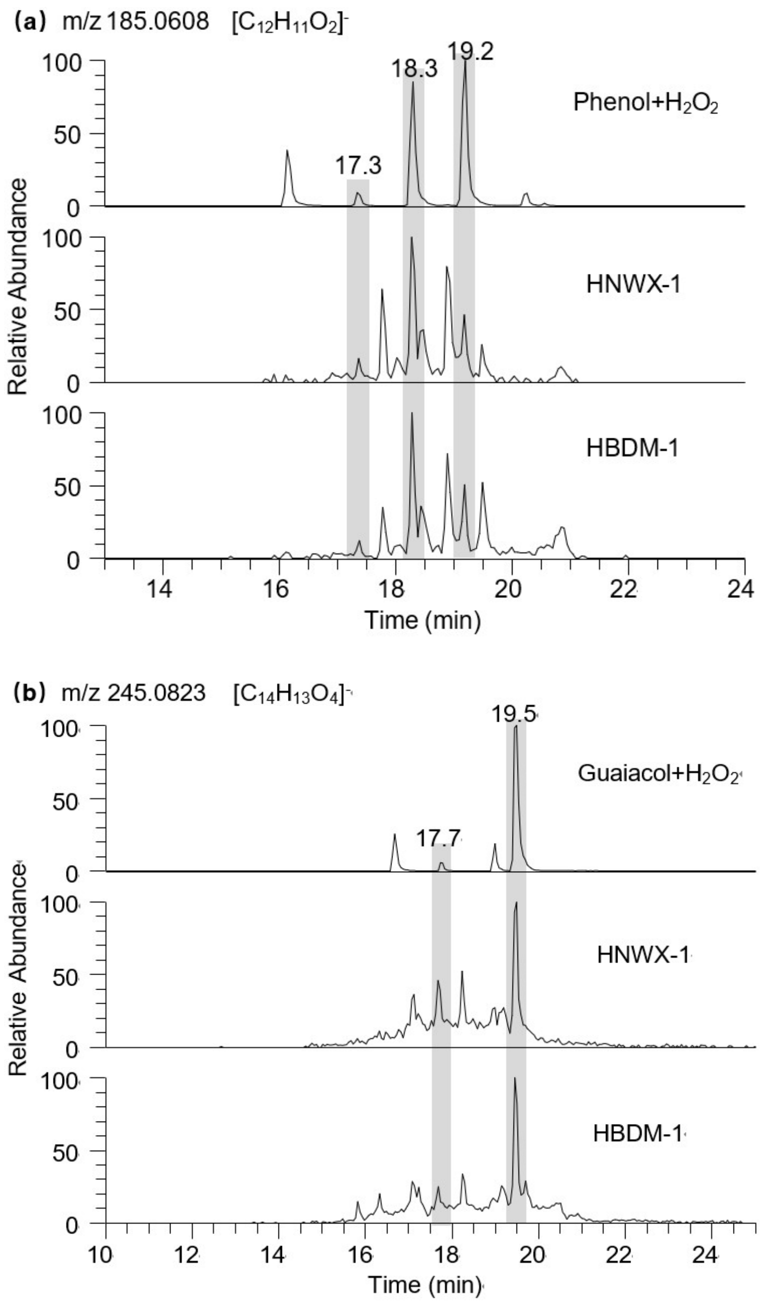

Figure 4. Extracted LC chromatograms of (a) m/z 185.0608 and (b) m/z 245.0823 in photochemical sample of phenol, HNWX-1, and HBDM-1, respectively.

The variation in peak abundance at unique retention times in the chromatogram could reflect the extent of evolution of WSOC molecules with accurate molecular weights, although no available standards were utilized for absolute quantification. The LC/ESI-HRMS monitors obviously changes in the molecular features of partial CHO species, i.e., photodegradation of low oxygenated compounds and formation of high oxygenated compounds. Table 1 lists the CHO compounds for which the LC peak intensities significantly increased and decreased after the 12-hour photolysis.

3.4.1 Photodegradation of low oxygenated compounds and formation of highly oxygenated compounds

As shown in Table 1, ion masses assigned with high unsaturated and low oxygenated species ($O/C < 0.5$) are prone to photodegradation, especially C_7-C_9 compounds (possible aromatic species), which intensity decreased by nearly one order of magnitude. For example, for m/z 123.0450 ($[C_7H_7O_2]^-$), as shown in Figure 5a, the peaks at RT 16.2 and 16.7 min in the LC chromatogram reduced in area by 95% after the 12-h irradiation. Using a standard it was verified that both peaks did not belong to guaiacol (peak at RT17.3 min), but they were also found within the products of guaiacol photooxidation, suggesting that they might be isomers of guaiacol or aromatic dihydric alcohol.

The phenolic dimers ($C_{12}H_{10}O_2$ and $C_{14}H_{14}O_4$) as described above also exhibited a decreasing tendency with almost complete disappearance after 12h direct photolysis. Other species with relatively high MW ($\geqslant$ 200Da) were also observed to be decomposed, including m/z 251.0564 ($[C_{12}H_{11}O_6]^-$), 313.0724 ($[C_{17}H_{13}O_6]^-$), and 329.0674 ($[C_{17}H_{13}O_7]^-$) (Figure S7), although their initial abundance was not very high.

Table 1. *M/Z* with significant changes upon 12-h photolysis analyzed by LC/ESI-HRMS.

Precursor (LC peak intensity decreases by >50%)			Product (LC peak intensity increases by >50%)		
Retention time, min	Measured *m/z*	Molecular formula	Retention time, min	Measured *m/z*	Molecular formula
16.2,16.7	123.04497	$C_7H_8O_2$	1.9	59.01362	$C_2H_4O_2$
13.9,14.5	129.05555	$C_6H_{10}O_3$	1.8	72.99291	$C_2H_2O_3$

14.6	131.07121	$C_6H_{12}O_3$	2.1	73.02928	$C_3H_6O_2$
14.6	133.02934	$C_8H_6O_2$	1.8	75.00856	$C_2H_4O_3$
15.9	135.04498	$C_8H_8O_2$	2.4	85.02930	$C_4H_6O_2$
13.7	137.02426	$C_7H_6O_3$	1.9, 4.4	87.04496	$C_4H_8O_2$
17.7	137.06063	$C_8H_{10}O_2$	1.9	88.98785	$C_2H_2O_4$
15.8	147.04504	$C_9H_8O_2$	1.9	89.02427	$C_3H_6O_3$
17.2	149.06062	$C_9H_{10}O_2$	2.2	99.00857	$C_4H_4O_3$
19.0	151.07634	$C_9H_{12}O_2$	2.5	129.01917	$C_5H_6O_4$
16.8	161.06068	$C_{10}H_{10}O_2$	2.0	145.01407	$C_5H_6O_5$
16.2	165.05559	$C_9H_{10}O_3$	1.9	147.02971	$C_5H_8O_5$
14.9	167.07129	$C_9H_{12}O_3$	14.9	155.03482	$C_7H_8O_4$
15.1	181.05048	$C_9H_{10}O_4$	15.1	169.01411	$C_7H_6O_5$
17.3	191.03498	$C_{10}H_8O_4$	16.4	183.02980	$C_8H_8O_5$
16.2	195.06622	$C_{10}H_{12}O_4$			
18.6	207.06635	$C_{11}H_{12}O_4$			

On the other hand, the solution acidity ([H^+]) of the particle extracts increase after the 12-hour photolysis, similar to the observation on the photooxidation of phenols (section 3.2) that resulted in the formation of oxygenated species. The solution acidity ([H^+]) normalized by WSOC concentration ([OC_{ws}]) was increased with a variation of $\Delta[H^+]/ [OC_{ws}]=(3.8\pm0.8)\times10^{-7}$ mol mgC^{-1}, suggesting the formation of new acidic substances.

The photochemical processing has led to an increased formation of low MW compounds (e.g., C_2-C_5 species), with a relatively high O/C. For example, the C_2 compounds, including $[C_2H_1O_3]^-$, $[C_2H_3O_3]^-$, $[C_2H_3O_2]^-$, and $[C_2H_1O_4]^-$ (Figure S8), which may correspond to glyoxylic acid, glycolic acid, acetic acid, and oxalic acid, respectively, were likely to be formed via oxidation pathway of several water-soluble molecules with photochemical reactivity (e.g., glyoxal (Carlton et al., 2007; Lim et al., 2010), methylglyoxal (Altieri et al., 2008; Lim et al., 2010), pyruvic acid (e.g. Grgic et al., 2010; Griffith et al., 2013; Reed Harris et al., 2014; Rapf et al., 2017; Eugene and

Guzman, 2017, Mekic et al., 2018; Mekic et al., 2019), phenols (Sun et al., 2010), etc). The presence of these highly oxygenated compounds that possibly contain acidic groups (e.g., –COOH and –OH) undoubtedly contributed to the increase of the solution acidity. Higher levels of other highly oxygenated species such as $[C_3H_5O_3]^-$, $[C_4H_7O_2]^-$, $[C_5H_5O_5]^-$ and $[C_5H_7O_5]^-$ were also observed (Figure S9).

To identify the impact of photolysis on the evolution of specific WSOC, the ions of $[C_7H_7O_n]^-$ in the HBDM-1 sample with significant variation were chosen as representative cases for description. The relative intensity of $[C_7H_7O_2]^-$ and $[C_7H_7O_3]^-$ decreased dramatically, while the intensities of $[C_7H_7O_4]^-$, $[C_7H_7O_5]^-$ and $[C_7H_7O_6]^-$ increased with the irradiation time (Figure 5 shows only the variation of $[C_7H_7O_2]^-$ and $[C_7H_7O_4]^-$ as an example). It seems reasonable that the possible hydroxylation of $[C_7H_7O_2]^-$ and $[C_7H_7O_3]^-$ might contribute to the formation of $[C_7H_7O_5]^-$ and $[C_7H_7O_6]^-$. Although we could not verify this hypothesis, the formed oxidized species undoubtedly have a high O/C ~~ratio~~ which highlights the possibility of this reaction pathway.

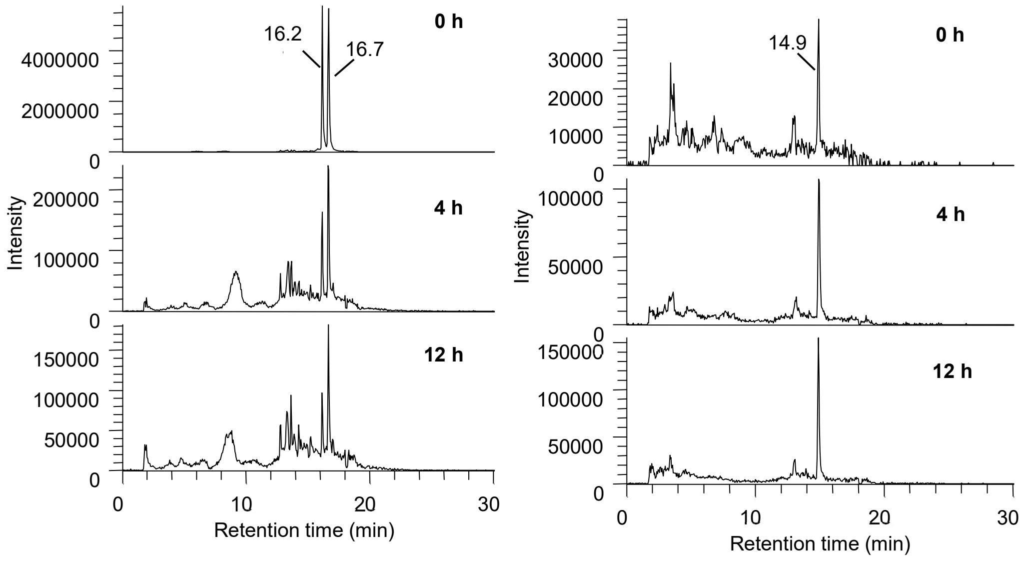

Figure 5. Extracted LC chromatograms from HBDM-1 of (a) $[C_7H_7O_2]^-$ and (b) $[C_7H_7O_4]^-$ at different photolytic stages of 0, 4, and 12 h.

3.4.2 Presentation of photochemically stable organic species

Some of the detected organic species seemed to exhibit a good photochemical stability, as their relative intensities only slightly decreased (<10%) after 12h light irradiation. The m/z 161.0454 ($[C_6H_9O_5]^-$) presented two prominent peaks at RT1.9 and 2.4 min (Figure S10). The peak at RT 2.4 min was further confirmed with a standard compound to be levoglucosan, a typical tracer of biomass burning aerosols with a high photochemical stability in atmospheric aerosols (Hu et al., 2013). The relatively good photochemical stability was also observed for some C_6 homolog compounds, such as $[C_6H_7O_6]^-$, $[C_6H_9O_6]^-$, and $[C_6H_{11}O_6]^-$. Some other oxygenated species, such as $[C_3H_3O_3]^-$, $[C_4H_5O_4]^-$, $[C_3H_3O_4]^-$, and $[C_4H_5O_5]^-$ remained relatively stable, as well.

Regarding the CHON compounds, only small variation of the chromatogram peaks, was observed for most of the detected species. In particular, several species with low O/C decreased by less than 30%, e.g., m/z 94.0297 ($[C_5H_4ON]^-$, RT 7.1 min), and 120.0453 ($[C_7H_6ON]^-$, RT12.2 min). Some compounds seem photochemically very stable as the variation of their peak intensities was less than 10 % upon light irradiation of the samples, e.g., m/z 118.0297 ($[C_7H_4ON]^-$, RT16.6 and 17.1 min), 146.0246 ($[C_8H_4O_2N]^-$, RT14.4 min), and 190.0510 ($[C_{10}H_8O_3N]^-$, RT17.8 min). However, the intensities of the ion masses with relatively higher degree of oxygenation was found to increase substantially (>50%), e.g., m/z 162.0195 ($[C_8H_4O_3N]^-$, RT 17.2 min), 198.0408 ($[C_8H_8O_5N]^-$, RT 18.0 min), and 242.1763 ($[C_{13}H_{24}O_3N]^-$, RT 17.9 min). The photochemical stability of some compounds may be ascribed to their low concentrations, or the light-shielding effect from other light-absorbing species.

Another intriguing finding was that different structural isomers with the same molecular mass might have exhibited different fates upon prolonged light irradiation of the samples. For example, the intensity of the peak at m/z 165.0405 ($[C_5H_9O_6]^-$) decreased when it was eluted at 4.9 min, but increased at RT 1.8 min, with the irradiation time (Figure S11). A simultaneous degradation and formation among isomers of some CHON ion masses upon prolonged light irradiation, was also observed, as was the case for the CHO compounds. For example, the m/z 108.0453 assigned to $[C_6H_6ON]^-$, might include hydroxy and amino groups on the phenyl ring to present three possible isomers (Figure S12). During photolytic processing, the intensity of the peak at RT 3.2 min increased dramatically, while there was a clear decreasing tendency of the peak intensity at RT 5.5 and 12.5 min, which was suggestive of possible isomerization among these isomers. Other ion masses that exhibited possible isomerization included m/z 122.0610 ($[C_7H_8ON]^-$), 132.0454 ($[C_8H_6ON]^-$), 134.0245 ($[C_7H_4O_2N]^-$), 136.0403 ($[C_7H_6O_2N]^-$), 138.0559 ($[C_7H_8O_2N]^-$), 144.0453 ($[C_9H_6ON]^-$), and 152.0352 ($[C_7H_6O_3N]^-$).

3.4.3 Comparison of time-profile mass spectra of CHO composition in WSOC extracts from WSBA samples

Since the LC method just separated a fraction of polar compounds, we tentatively utilized the change of HRMS to gain more comprehensive information about the WSOC evolution. We compared the time-profile (0, 4, and 12h) mass spectra with each other, based on the assumption of same interference from inorganic species, and the good reproducibility and stability for Orbitrap MS operated under the same instrumental parameters (the RSD of TIC intensity within 5%). It is well known that ESI mass spectral abundances are influenced by the solution composition, concentration of analytes and instrumental factors (Bateman et al., 2011); hence, it is quite challenging to directly quantify the absolute concentration levels of the complex mixtures. Despite that, the photochemical degradation of WSOC compounds and corresponding formation of organic compounds can be well described by the variation of signal intensity from mass spectrometry. The average O/C and H/C ~~ratios~~ for CHO compounds were from 0.38±0.02 to 0.44±0.02 and 1.24±0.03 to 1.26±0.01, respectively, as the irradiation time extended ~~up~~ from 0 to 12h. The comparison of these time-profile mass spectra indicates that the 12-hour photolysis resulted in a significant reduction of 28±11% in the total ion abundance (S/N). Since the photolysis induced changes in abundance for most of the CHO compounds, we also calculated the intensity (S/N)-weighted average O/C (O/C_w) and H/C (H/C_w) (Bateman et al., 2011; Romonosky et al., 2015) with values ranging from 0.45±0.03 to 0.53±0.06 and from 1.32±0.09 to 1.40±0.11, respectively. After the 12-h photolysis, both average H/C and H/C_w values slightly increased, compared to the samples prior to irradiation, however, both average O/C and O/C_w values have increased more distinctly, indicating an elevation in oxidation degree of bulk extract composition. This phenomenon could be partly reflected on the LC-HRMS observation, i.e. formation of highly oxygenated species and the consumption of low oxygenated compounds. In our previous study, the UV-VIS measurements revealed that the 12-h photochemical evolution leads to a modification of absorptive properties for WSBA extracts (e.g., photo-bleaching at wavelengths below 380nm and photo-enhancement above 380nm) (Cai et al., 2018), which might be partially linked to present findings about molecular functionalization, e.g., hydroxylation facilitating a red shift for light absorbing wavelengths.

4 CONCLUSIONS

This study was focused on the effect of direct photolysis on the molecular composition of actual WSOC extracted from field straw-burning aerosol. The phenol dimer (m/z 185.0608) and guaiacol dimer (m/z 245.0823), or their isomers generated from laboratory aqueous-phase photooxidation of phenol and guaiacol were also observed in present field WSBA samples, suggesting that the aqueous phase reaction might contribute to the formation of emitted biomass burning aerosols. The laboratory observation on aqueous photochemistry of phenols indicated that those phenolic compounds in real biomass burning aerosols would likely have potential to experience ~~the~~ similar evolution to form various oxygenated compounds under relevant atmospheric water conditions. The direct photolysis on the molecular composition of WSOC extracts from WSBA samples were performed to gain more insight into the evolution of aerosol composition. Because the extract composition was very complex, the techniques (ESI-HRMS and LC/ESI-HRMS) used in this study, although advanced still had limitations in monitoring the modification of molecular composition, especially for determining the potential formation of compounds present at low concentrations or compounds that were poorly ionized. However, a series of polar molecules were identified that changed their molecular composition via photochemical aging. In particular, the degradation of low oxygenated compounds with strong photochemical reactivity and the formation of high oxygenated compounds might directly result in an increasing O/C in WSOC composition, which was likely linked to the modification of light-absorbing characteristics for extracts in previous study. This finding indicates that the water soluble organic fraction of field combustion-derived aerosols has the potential to form more oxidized organic matter, which might ~~partly account for~~contribute to the highly oxygenated nature of atmospheric organic aerosols. Further studies focused on the photochemical evolution of WSOC composition will be performed in the future, including enlarging measurements on compound species (e.g., applying positive ESI-HRMS), identifying biomarkers and evaluating their role in photochemical processes.

AUTHOR CONTRIBUTION

Jing Cai and Zhiqiang Yu designed the experiments, and Jing Cai and Xiangying Zeng carried them out. Guorui Zhi provided the straw-burning aerosol samples, Zhiqiang Yu and Sasho Gligorovski helped to perform the analysis of light irradiation and editing the manuscript. Guoying Sheng, Xinming Wang and Ping'an Peng provided some technical consultations about organic chemistry. Jing Cai prepared the manuscript with contributions from all co-authors.

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