Reply to RC1

*In the following, the referee's comments are reproduced (black) along with our replies (blue)
and changes made to the text (red) in the revised manuscript.*

General statement:

Eger et al. present measurements of ClNO2, HCl, SO2, O3, and NO2 mixing ratios in the
Eastern Mediterranean Sea around the Arabian Peninsula during summer 2017. The data
inform about the conversion of N2O5 to ClNO2 in this environment and are a welcome
addition to the literature as there are few such measurements outside North America, China
and Europe. The analysis is thorough and shows that the ClNO2 production from NO3
(epsilon) is small, which is not surprising considering the warm temperatures that shift the
equilibrium away from N2O5 chemistry.
We thank the referee for the positive evaluation of our manuscript and the useful comments
and suggestions. We modified the manuscript according to the comments listed below.

General comments:

The data set is broken up into sections by region (Gulf of Oman, Gulf of Aden, Red Sea, etc.)
and presented as (nocturnal) averages, median, and maxima (Table 1 + Figure 2). I felt that
this wasn't the most accessible way to present the data and added confusion. Examples are
the averaged nocturnal locations shown in Figs 2 and 5. It would have been more transparent
to present a continuous trace of ship locations color-coded by ClNO2 mixing ratios (and split
the figure up into part a = leg 1 and part b = leg 2).
In line with these suggestions we have redrawn Fig. 2 which now shows a continuous trace of
nocturnal data points colour-coded by $ClNO_2$ mixing ratios and split up into first and second
leg. To avoid excessive overlap of data points we use 1-hour averages instead of the original
5 min data. The former Fig. 2 was shifted to the supplement (now Fig. S6) and the manuscript
text has been modified:
"Maximum $ClNO_2$ mixing ratios observed during each night ranged from the limit of detection
to 586 pptv (see Fig. S6 for details). Figure 2 shows 1-hour averaged $ClNO_2$ mixing ratios
along the ship track during (a) first and (b) second leg. Text boxes indicate the median night-
time mixing ratios of O_3, HCl, NO_2 and SO_2 for the different regions where data from the first
and second leg datasets have been combined."
We also added a similar plot color-coded by ε (complemental to Fig. 5) to the supplement to
provide additional information. A reference was added to the text.
"[…] (for a more detailed plot with 1-hour averaged data points see Fig. S13)."

After all, averages can be skewed by plumes.
The nocturnal averages of ε in Fig. 5 are median values and thus less influenced by single plumes than mean values would be.

5

I also felt that the data were over-interpreted since changes are interpreted as regional differences rather than temporal ones. Imo, it would have perhaps been more informative to stay put in one or two places for some time for that reason, but that was perhaps outside the control of authors.

10 The aim of the AQABA campaign was to achieve a large spatial coverage around the Arabian Peninsula (within a reasonable time period of 2 months) as the whole region is severely understudied. It was thus not intended and out of the control of the authors to stay longer in some area to increase statistics. Despite high temporal variability in observed $ClNO_2$ mixing ratios within one region, we still think that the separation into different regions is a useful way

15 to present the large amount of data and is not unjustified with respect to the different air mass characteristics (e.g. NO_X levels).

Overall, the manuscript is suitable for ACP. However, the manuscript is on the long side and could (and perhaps should) be condensed.

20 The analysis of this large dataset covering a two-month campaign is quite complex and the paper is necessarily a bit on the long side. To keep it as short as possible we already had moved information (e.g. details of the calculation of ε and corrections to the aerosol particle surface area concentration) to the supplement and felt we had the right balance of information in the main manuscript and the supporting information. We have not been able to identify

25 further text sections we could easily move to the supplement without perturbing the basic structure of the manuscript. However, we shifted Fig. 7 (calculation of f via first method, now Fig. S14) to reduce the amount of figures.

Some figures are missing axis labels (Latitude, longitude, day, month, etc.).
30 We added labels to the figures where they were missing, see specific comments below.

There are also few typos in equations that need to be corrected.
Typos in equations have been corrected, see specific comments below.

35 The date formatting does not adhere to ACP standards, e.g., 25 July 2007 (dd month yyyy), not 25.7.
We changed the date formatting throughout the manuscript, see specific comments below.

Specific comments:

Page 2, reaction (6). There is also a minor channel producing NO2.

We added the reaction for the minor production channel and changed the labelling of the two linked reactions. We now write:

$NO_3 + hv \qquad\qquad \rightarrow \qquad\qquad NO + O_2$ (R6a)

$NO_3 + hv \qquad\qquad \rightarrow \qquad\qquad NO_2 + O$ (R6b)

Page 4 line 29. You would get two time series, one for m/z 208 and another for m/z 210. Please specify how you used two masses (added them, or averaged them) or did you chose one over the other?

We chose m/z 208 for its higher S/N-ratio to calculate the $ClNO_2$ mixing ratios. The mean ratio of m/z 208 to m/z 210 for the whole campaign was 3.08 (which is very close to the theoretical value of 3.13) with a correlation coefficient of R^2 = 0.96. We added the corresponding plot to the supplement and modified the text with the following:

"We chose the signal at m/z 208 for its higher signal-to-noise (S/N) ratio to calculate the $ClNO_2$ mixing ratios reported. For the whole campaign dataset, the ratio between m/z 208 and m/z 210 was 3.08 (R^2 = 0.96, see Fig. S1) which is very close to the expected value of 3.13 derived from the natural abundance of the ^{35}Cl and ^{37}Cl isotopes, indicating no significant interferences at either of the two m/z."

Note that you can get a IBr- at 208.

According to Liao et al. (2011), $HO^{81}Br$ can also be detected as $I^{81}Br^-$, possibly interfering with $IClNO_2^-$ at m/z 208. However, this is a minor channel compared with the formation of $IHO^{81}Br^-$ (m/z 223) and, as stated above, the correlation between m/z 208 and m/z 210 was very good, indicating that we did not detect any significant interference.

Liao, J., et al. "A comparison of Arctic BrO measurements by chemical ionization mass spectrometry and long path-differential optical absorption spectroscopy." *Journal of Geophysical Research: Atmospheres* 116.D14 (2011).

Line 31 – here, you give one sensitivity. Shouldn't the sensitivity at m/z 210 be ~1/3 that of m/z 208?

Yes, that is true. We added the sensitivity for m/z 210 to the text, which is 0.20 Hz pptv^{-1}.

"$I \cdot ClNO_2^-$ is more specific than ICl^- (m/z 162 and 164) and has a lower background signal, providing a sensitivity of 0.61 Hz pptv^{-1} per 10^6 Hz of I^- at m/z 208 (and 0.20 Hz pptv^{-1} at m/z 210), a limit of detection (LOD) (2σ, 5 min) of 12 pptv and a total measurement uncertainty of 30 % ± 6 pptv."

Page 5- line 2. Again, how did you use two m/z values to get one mixing ratio?

We only used *m/z* 188 to calculate the HCl mixing ratio, as *m/z* 190 suffers from a high background signal and an interference from I·HNO$_3^-$. We added the sensitivity for *m/z* 190 to the text, which is 0.05 Hz pptv^{-1}.

"HCl was observed as I(CN)Cl$^-$ (*m/z* 188 and 190) (Eger et al., 2019) with a sensitivity of 0.17 Hz pptv^{-1} per 10^6 Hz of I$^-$ at *m/z* 188 (and 0.05 Hz pptv^{-1} at *m/z* 190), a detection limit of 98 pptv and a total measurement uncertainty of 20 % ± 72 pptv. As *m/z* 190 suffers from known interferences (e.g. I·HNO$_3^-$) and has a lower S/N ratio, we used *m/z* 188 to calculate the HCl mixing ratios reported."

Page 5. Please comment if the stack emissions truly interfered with ClNO2 measurement by CIMS, or if the data were filtered simply as a precaution.

The datasets of all instruments sampling from the common inlet were filtered for our own stack emissions as a precaution because these fresh emissions (containing large amounts of particles, NO$_x$, hydrocarbons, black carbon, soot etc.) superimpose with the measured air masses and can potentially bias the results. Data where NO is above background level (like in our own ship's plume) is excluded from the calculation of ε, anyway. We amended the text:

"All datasets were filtered prior to analysis for periods where the measurements were contaminated by stack emissions to avoid a potential bias in the results."

Page 6 line 12 "modified" how? Was it equipped with a photolytic converter?

We agree that the word "modified" adds confusion, so we decided to remove it and to add a reference instead (Li et al., 2015), describing the instrument with its modifications.

"NO and NO$_2$ were measured by a chemiluminescence detector (CLD 790 SR, ECO Physics, Duernten, Switzerland) (Fontijn et al., 1970; Li et al., 2015)."

Page 6 lines 18-19. Meusel et al. 2016 state that J values were not corrected for up-welling UV radiation. This should also be stated here since it biases the J values low.

That is right, the J-values were not corrected for upwelling UV radiation, which is included in the overall uncertainty. We added a note to the text:

"J-values were not corrected for upwelling UV radiation and are estimated to have an overall uncertainty of ≈ 10 %."

Page 9 equation (2). [O3] also changes over time.

In our simple calculation of ε we assume that [O$_3$] does not change over time, as already mentioned in the text. The relative decrease in [O$_3$] over time is usually negligible (< 10 %), given the total uncertainty of the calculation, whereas the relative increase in [NO$_2$] from time t_0 to t can be large and has to be accounted for. We added a line to emphasise this:

"In this calculation we assume that [NO$_2$] changes over time but [O$_3$] stays constant in good approximation."

Page 10 "Boundary layer height of 1000 m". That seems high for the marine boundary layer. Is there evidence to corroborate such a high mixing height?

We used an estimated boundary layer height only to give an example of potential rates of HNO_3 loss, which would impact on our calculation of the reaction time via equation (5). We also state that we relax the criterion for a match between calculated time and time elapsed since the beginning of the night as such effects, which depend i.a. on the boundary layer height and the HNO_3 to NO_2 ratio, are rather uncertain. We have deleted the reference to boundary layer height as it was NOT used in correction, and may have been misleading.

Page 11 equation (8). Please define Keq. Should it be capitalized?

Yes, K_{eq} should be capitalized. We added a definition to the text.

"[…] where A is the particle surface area concentration, $\bar{c}$ is the mean molecular velocity of N_2O_5 ((24400 ± 160) cm s^{-1} during AQABA) and $K_{eq} = \frac{[N_2O_5]}{[NO_2]\,[NO_3]} = 2.8 \times 10^{-27}$ (T/300)$^{-0.6}$ exp(11000/T) cm^3 molecule^{-1} (IUPAC, 2019) is the temperature-dependent equilibrium constant (Reactions R4 and R5)."

Page 13 Please check equation 11 – does not look right.

Line 2 – "k" does not appear in equation 11; perhaps it should say "B" rather than "Bk"?

The equation was corrected (brackets were placed the wrong way) and parameters were renamed for better readability.

$$\gamma = B \times k \times \left(1 - \left(\left(a \times \frac{[H_2O(l)]}{[NO_3^-]}\right) + 1 + \left(b \times \frac{[Cl^-]}{[NO_3^-]}\right)\right)^{-1}\right) \tag{11}$$

where $B = 3.2 \times 10^{-8}$ s, $k = 1.15 \times 10^6$ - 1.15×10^6 exp(-0.13 [H$_2$O(l)]) s^{-1} is the rate constant for the reaction $N_2O_5(aq) + H_2O(l)$, $a = 0.06$ denotes the ratio of rate constants for reactions $H_2NO_3^+(aq) + H_2O(l)$ and $H_2NO_3^+(aq) + NO_3^-(aq)$ and $b = 29$ denotes the ratio of rate constants for reactions $H_2NO_3^+(aq) + Cl^-$ and $H_2NO_3^+(aq) + NO_3^-(aq)$.

Line 10. 14+/-14% and throughout the paper. Since the % operator applies to both 14 and 14, please add brackets (14+/14)%. Otherwise, it reads as a percent error, i.e., is interpreted as 14+/-2. This is repeated throughout the paper (e.g., line 21, 13+/10% could be 13+/-1.3). Same goes for units when uncertainties are given (e.g., line 31, 25-35 °C)

This was corrected throughout the manuscript.

Page 15 line 15 – punctuation error. Note that there are others like this throughout the paper.

Punctuation errors were corrected throughout the manuscript.

Equation 15 is incorrect.
We corrected the typo in the equation:

$$p\mathrm{OH}_{\mathrm{O3}} = \frac{2\, J_{\mathrm{O1D}}\, [\mathrm{O_3}] \,\times\, k_{\mathrm{H2O}}\, [\mathrm{H_2O}]}{k_{\mathrm{H2O}}\, [\mathrm{H_2O}] + k_{\mathrm{N2}}\, [\mathrm{N_2}] + k_{\mathrm{O2}}\, [\mathrm{O_2}]}$$

Page 19. Many references are missing doi's.
DOIs have been added.

Figure 2 and Figure 3. Please label all axes for clarity.
Missing labels have been added to the figures. Dates have been modified (e.g. "Day in July").

Figure 4 A lot of the variability may be due to not having enough data. Consider longer averages (1 hr, 90 min, or 2 hr) for the ClNO2 data.
The variability is mainly due to a mixture of high atmospheric variability and a limited number of days we spent in one region. Changing the averaging interval to e.g. 1 hour did not significantly change the shape of the curves.

Figure 7. Are you sure if this analysis is valid? It is possible that changes in NOz and ClNO2 are due to shifting air mass.
(Note: The former Fig. 7 was shifted to the supplement and is now Fig. S14, see above.)
For this analysis we assume (as stated in the text) that we sample a homogeneous air mass (indicated by wind direction, T, RH etc.), i.e. changes in NO_z and $ClNO_2$ are not caused by a change of the air mass within the period of observation. As this requirement was rarely fulfilled, we could only analyse the four different episodes mentioned in the text.

Figure 8. Are these total chloride and total sodium concentrations, or from certain size fractions only? (state in caption)
(Note: The former Fig. 8 is now Fig. 7.)
These are PM_1 data only (AMS measurements, see Sect. 2.5). We added the information to the caption:
"Co-variance between mixing ratios of SO_2, NO_2 and HCl and particulate chloride depletion (calculated from Eq. 12) illustrated by the difference in Cl^- and Na^+ (PM_1) measured."

Figures 9 and 10. I think it's important to point out here that the authors only consider selected sources of radicals (ClNO2 photolysis and O1D+H2O). Some important ones are omitted (such as Cl2 and HONO photolysis and HO2+NO).
We added text to the beginning of section 3.6:
"Other potential Cl sources (e.g. Cl_2 photolysis) are not considered here as we do not have experimental data to quantify their impact."

We also added a sentence to the paragraph where we describe the calculation of $P_{OH}(O_3)$:
"As we do not consider other OH production channels (e.g. photolysis of HONO or HO_2 + NO), which can be of importance under more polluted conditions, pOH_{O3} represents a lower limit of

5   pOH."

The labelling and caption of the former Fig. 9 (now Fig. 8) has been modified to emphasise that only exclusive channels were considered:

"Time series of $ClNO_2$ mixing ratios, J_{ClNO2} photolysis rates and production of Cl-radicals from $ClNO_2$ photolysis (pCl_{ClNO2}) and OH-radicals from O_3 photolysis in the presence of H_2O

10   (pOH_{O3}) for two consecutive nights in the Gulf of Oman."

Figure 10 I wouldn't lump HCl+OH->H2O+Cl (a conversion of one radical to another) in with OH and Cl production from O3 and ClNO2 photolysis (which generate radicals from stable molecules).

15   In former Fig. 10 (now Fig. 9) we only compare the relative contribution of HCl + OH and $ClNO_2$ + hv to Cl radical formation. Although transformation of OH into Cl does not change the overall radical budget nor the RO_2 budget, the relative oxidation rates of several VOCs will be modified as Cl reacts much faster with some of them than OH does.

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Reply to RC2

In the following, the referee's comments are reproduced (black) along with our replies (blue)
and changes made to the text (red) in the revised manuscript.

General statement:

This paper (Shipborne measurements of ClNO2 in the Mediterranean Sea and around the Arabian Peninsula during summer) reports observations of ClNO2, NO3/N2O5, HCl, particle composition and other parameters made during a cruise in the Mediterranean Sea, Red Sea and Persian Gulf. This is a severely understudied region in terms of atmospheric chemistry and, as such, the dataset presented here fills a significant gap. The paper is well laid out, the figures and tables clear, and the analysis of the data is interesting and thorough. I only have a few minor observations, but other than that, I recommend publications on ACP.
We thank the referee for the positive evaluation of our manuscript and the useful comments and suggestions. We modified the manuscript according to the comments listed below.

General comments:

I find the analysis in Section 3.4 a bit confused. First of all, a little introduction explaining how the factors influencing ClNO2 production efficiency are going to be evaluated in this section would be useful in order to follow the discussion.
We added an introductive sentence to the section:
"In the following we calculate f and γ from our measurements, compare the values with the literature and quantify the contributions of k_{het} and k_{dir} to the overall NO_3 loss rate."

Second, the values of f calculated with Eq 9 and with Eq 10 are significantly different, but this discrepancy is not really explained or discussed.
With Eq. 10 we calculated median values of f for each region, based on all available data, and listed them in Table 2 (here we added a note: "[a] Calculated from Eq. (10)"). In contrast, Eq. 9 could only be applied to four specific time periods listed in the text, when we sampled a homogeneous air mass. There is no reason to expect a perfect agreement as these four values calculated with Eq. 9 are only snapshots within different regions.

It is also not clear if the value for the Gulf of Oman is 0.6 (page 11, line 29) or 0.84 (page 12, line 3).
These are two values for two different time periods, both in the Gulf of Oman (first one shown in Fig. S14 (formerly Fig. 7), second one not shown). We added a date to the first event to make this clear.

"[…] as illustrated in Fig. S14 for data obtained in the Gulf of Oman (25–26 July 2017) for which $f = 0.60 \pm 0.04$."

5 When it comes to f, the main issue is the availability of particulate chloride. In general, it seems (page 12, lines 20-25) that the authors are focusing on fine particles, while I would expect sea salt to be a dominant source of chloride in the open sea. It may be true that the surface area of sea salt is smaller but the ClNO2 yield is higher, as the authors themselves acknowledge on page 13. Therefore neglecting sea salt in the calculation of f may not be
10 appropriate and could possibly lead to a bias in the results of the analysis.
We agree that sea salt may contribute to $ClNO_2$ formation (due to an f close to 1). However, the average contribution from the coarse mode during the campaign was only 14 % (as stated in the text) and f is between 0.5 and 1 for fine mode particles as well, so the fractional contribution of coarse mode sea salt to $ClNO_2$ formation is generally low (as most of the
15 coarse mode particles were dust). This is discussed in detail towards the end of Sect. 3.4.
We note that, the calculation of $\varepsilon = \left(\frac{[ClNO_2]}{[NO_3]_{int}} \right)$ is in any case independent of k_{het} (and thus independent of the contributions of the coarse/fine mode).

Finally the statement on page 13 line 30 about the importance of kdir, i.e. the direct NO3 loss,
20 seems to be in contrast with the last lines of the section. I am afraid it is not enough to refer to a future publication, given that a significant part of the analysis stands on the assumption that the direct losses of NO3 dominate over the indirect losses. At least a summary of the steady state analysis mentioned here should be given to support the statements about kdir.
The last lines of the section state that much of the reactivity could neither be attributed to k_{het}
25 nor to measured VOCs but to unidentified compounds (also contributing to k_{dir}), so they are not in contrast with the statement that k_{dir} is more important than k_{het}.
The result that $k_{dir} \gg k_{het}$ is also derived from measurement data via $\varepsilon = f \left(\frac{k_{het}}{k_{het}+k_{dir}} \right)$.
To make this clear, we modified the text:
"However, for a large fraction of each night NO_3 was below the detection limit (ca. 5 pptv)
30 despite a high production rate (large mixing ratios of NO_2 and O_3). A steady-state analysis of NO_3 production and loss indicated a high total reactivity which could not be attributed to measured trace gases (k_{dir}) or heterogeneous losses of N_2O_5 (k_{het}). A detailed analysis of the NO_3 lifetime and the role of VOCs is beyond the scope of the present manuscript and will be described in detail in a separate publication."

35

Specific comments:

page 1, line 30: capitalize "Earth"

Done.

"As the Arabian Gulf already suffers from some of the most polluted air on Earth […]"

page 5, line 2: I am not sure I follow the ion chemistry from HCl to I(CN)Cl-. Where is the CN group coming from? Please provide more information or add the relevant reference.

Detection of HCl involves $I(CN)_2^-$ primary ions. A reference (Eger et al., 2019) was added.

"HCl was observed as $I(CN)Cl^-$ (m/z 188 and 190) (Eger et al., 2019) with a sensitivity of 0.17 Hz $pptv^{-1}$ per 10^6 Hz of I^- at m/z 188 (and 0.05 Hz $pptv^{-1}$ at m/z 190), a detection limit of 98 pptv and a total measurement uncertainty of 20 % ± 72 pptv."

page 5, line 7 and 12: can you provide more information on the purpose of the IMR bypass?

The bypass in front of the IMR (1 slm) was used in order to improve the transmission of $CH_3C(O)O_2$ radicals from the thermal decomposition of PAN in the heated inlet (Eger et al. (2019). As this is not relevant for $ClNO_2$, HCl or SO_2 detection, we removed the sentence.

And it is not clear to me how 50 cm of a 1/8 inch tube reduces the pressure in a 3 m long inlet.

The pressure is both reduced by a bypass flow (5 slm) and by the mentioned (coiled) piece of 1/8 inch tube (which is reducing the pressure in the inlet line via energy dissipation to the walls). We added the missing information to the text (and changed units from inch to mm):

"To avoid condensation of water in the inlet lines in the containers, the pressure in the sampling line was reduced to ≈ 700–800 mbar with a bypass flow of ≈ 5 slm and by including an additional ≈ 50 cm long (coiled) piece of 3.18 mm (OD) PFA tubing."

page 6, line 10: do you mean NO3?

NO_2 is correct.

page 10: can you specify which of the methods explained in the supplement is being used as default in the paper discussion and in Figure 6? I am guessing method B but it should be stated.

Method C is used as default as stated in the supplement. We wanted to avoid confusion by referring to a method that is only mentioned in the supplement, since the data reduction is also (briefly) outlined in the manuscript itself. Nevertheless, we added a note with a reference to the supplement to Sect. 3.2.

"The data reduction is described in more detail in the supplement (all the data shown in the manuscript corresponds to the application of method C), where the sensitivity of ε to these limitations and additional constraints is discussed."

equation 8: I think you need to explain the keq[NO2] part of the equation and how it is related to [N2O5].

We added a definition to the text:

"[…] where A is the particle surface area concentration, $\bar{c}$ is the mean molecular velocity of N_2O_5 ((24,400 ± 160) cm s^{-1} during AQABA) and $K_{eq} = \frac{[N2O5]}{[NO2]\,[NO3]} = 2.8 \times 10^{-27}$ (T/300)$^{-0.6}$ exp(11000/T) cm^3 molecule^{-1} (IUPAC, 2019) is the temperature-dependent equilibrium constant (Reactions R4 and R5)."

At several points in the paper the notation IClNO2- (or similar) is used for the masses measured by CIMS. But IClNO2 is a cluster not a molecule, so it should be more correctly indicated as I.ClNO2-. The same for other ions mentioned throughout the paper.

This was corrected throughout the manuscript. We now write $I \cdot ClNO_2^-$ and $I \cdot HNO_3^-$.

Shipborne measurements of ClNO₂ in the Mediterranean Sea and around the Arabian Peninsula during summer

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Abstract. Shipborne measurements of nitryl chloride (ClNO₂), hydrogen chloride (HCl) and sulphur dioxide (SO₂) were made during the AQABA (Air Quality and climate change in the Arabian BAsin) ship campaign in summer 2017. The dataset includes measurements over the Mediterranean Sea, the Suez Canal, the Red Sea, the Gulf of Aden, the Arabian Sea, the Gulf of Oman and the Arabian Gulf (also known as Persian Gulf) with observed ClNO₂ mixing ratios ranging from the limit of detection to ≈ 600 pptv. We examined the regional variability in the generation of ClNO₂ via the uptake of dinitrogen pentoxide (N₂O₅) to Cl-containing aerosol and its importance for Cl-atom generation in a marine boundary layer under the (variable) influence of emissions from shipping and oil industry. The yield of ClNO₂ formation per NO₃ radical generated was generally low (median of ≈ 1–5 % depending on the region), mainly as a result of gas-phase loss of NO₃ dominating over heterogeneous loss of N₂O₅, the latter being disfavoured by the high temperatures found throughout the campaign. The contributions of ClNO₂ photolysis and OH-induced HCl oxidation to Cl-radical formation were derived and their relative contributions over the diel cycle compared. The results indicate that over the northern Red Sea, the Gulf of Suez and the Gulf of Oman the formation of Cl-atoms will enhance the oxidation rates of some VOCs, especially in the early morning.

1 Introduction

The AQABA (Air Quality and climate change in the Arabian BAsin) campaign was designed to study air quality and climate in a region (Eastern Mediterranean and Middle East) that is likely to be heavily impacted by future climate change with increasing frequency and intensity of droughts, heatwaves and associated Aeolian dust and pollution emissions (Lelieveld et al., 2012). As the Arabian Gulf already suffers from some of the most polluted air on Earth with O₃ levels regularly greater
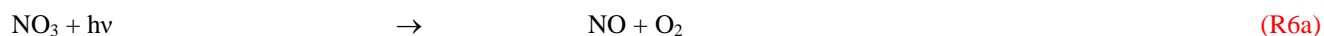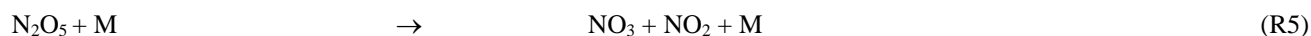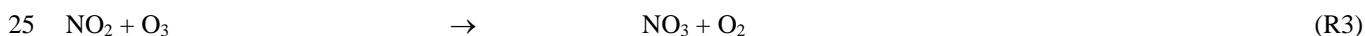
12

than 100 ppbv (Lelieveld et al., 2009), one aspect of the campaign was to investigate the factors that contribute to high levels of air pollution in the region. This includes the impact of reactive chlorine chemistry resulting from the interactions of pollutant emissions from ships and petrochemical activity with sea-salt, under conditions influenced by intense photochemistry and high temperatures during summer.

5 The heterogeneous uptake of gaseous N_2O_5 to the aerosol phase represents an important atmospheric sink for NO_x (NO + NO_2) via conversion to nitric acid (HNO_3), which is efficiently removed from the boundary layer via deposition (Lelieveld and Crutzen, 1990; Dentener and Crutzen, 1993; Macintyre and Evans, 2010). In the presence of aerosol chloride, nitryl chloride ($ClNO_2$) can also be formed along with HNO_3 (NO_3^-) as shown in Reaction (R1) (Finlayson-Pitts et al., 1989; Behnke et al., 1997). $ClNO_2$ has a lifetime of more than 30 hours in the nocturnal marine boundary layer (Osthoff et al.,

10 2008) but is rapidly photolysed after sunrise (Reaction R2), releasing nitrogen dioxide (NO_2) and chlorine atoms.

N_2O_5 + (H_2O or Cl^-)	$\rightarrow$	(2-f) NO_3^- + f $ClNO_2$, $0 \leq f \leq 1$	(R1)
$ClNO_2$ + hv	$\rightarrow$	Cl + NO_2	(R2)
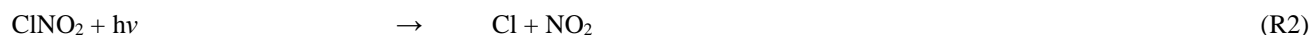
The formation of $ClNO_2$ can have a significant impact on regional NO_x cycling and radical chemistry especially in the polluted coastal and marine boundary layer (Simon et al., 2009; Riedel et al., 2014; Sarwar et al., 2014). The Cl-atoms

15 formed in Reaction (R2) can enhance oxidation rates of several volatile organic compounds (VOCs) especially during early morning hours (Phillips et al., 2012; Riedel et al., 2012a; Young et al., 2012) thus contributing to photochemical ozone production (Simon et al., 2009; Riedel et al., 2014; Sarwar et al., 2014; Faxon et al., 2015; Wang et al., 2019).

The chemical processes involved in the formation of $ClNO_2$ are complex and, as outlined in Fig. 1, involve the sequential oxidation of NO_x to N_2O_5 via NO_3 (Reactions R3–R5). During the day NO_3 is rapidly photolysed via Reaction (R6) or reacts

20 with nitrogen oxide (NO) via Reaction (R7) so that N_2O_5 formation is supressed. The heterogeneous reaction of N_2O_5 with particles is thus, to a good approximation, limited to the night-time. The equilibrium between NO_3 and N_2O_5 (Reactions R4 and R5) is strongly temperature-dependent, with N_2O_5 formation favoured by high NO_2 mixing ratios and low temperatures. NO_3 can also react with volatile organic compounds (VOCs) (R8) forming e.g. alkyl nitrates, which also reduces the rate of formation of N_2O_5.

25 NO_2 + O_3	$\rightarrow$	NO_3 + O_2	(R3)
NO_3 + NO_2	$\rightarrow$	N_2O_5	(R4)
N_2O_5 + M	$\rightarrow$	NO_3 + NO_2 + M	(R5)
NO_3 + hv	$\rightarrow$	NO + O_2	(R6a)
NO_3 + hv	$\rightarrow$	NO_2 + O	(R6b)
30 NO_3 + NO	$\rightarrow$	2 NO_2	(R7)
NO_3 + VOCs	$\rightarrow$	products	(R8)

The N_2O_5 loss rate via heterogeneous uptake to particles is described by Eq. (1) where $\bar{c}$ is the average molecular velocity of N_2O_5, A is the particle surface area concentration and γ is the uptake coefficient.

$$\frac{\mathrm{d}[N_2O_5]}{\mathrm{d}t} = -0.25\,\bar{c}\,\gamma\,A\,[N_2O_5] \tag{1}$$

The uptake coefficient, γ, has been characterised in several laboratory investigations (see Bertram and Thornton (2009); Chang et al. (2011); Ammann et al. (2013) for summaries) and in numerous field studies where it has been found to be highly variable (between 5×10^{-4} and 0.11) and dependent on temperature, relative humidity (RH) and aerosol composition (Brown et al., 2006; Bertram et al., 2009; Brown et al., 2009; Riedel et al., 2012b; Wagner et al., 2013; Morgan et al., 2015; Brown et al., 2016; Phillips et al., 2016). A value of ≈ 0.03 has been derived from measurements in the polluted marine environment (Aldener et al., 2006).

The $ClNO_2$ yield, f, which controls the relative formation rates of NO_3^- and $ClNO_2$ in Reaction (R1), is determined by the $[Cl^-]$ to $[H_2O]$ ratio in the aerosol phase (Behnke et al., 1997; Bertram and Thornton, 2009; Ammann et al., 2013), and can vary between zero to unity (Thornton et al., 2010; Wagner et al., 2012; Riedel et al., 2013; Phillips et al., 2016; Wang et al., 2016; McDuffie et al., 2018b). In Fig. 1 we introduce the $ClNO_2$ production efficiency ε, which is the yield of $ClNO_2$ per NO_3 molecule formed in Reaction (R3) and will be discussed in detail in Sect. 3.2.

The established method to measure atmospheric $ClNO_2$ mixing ratios from a few tens of pptv (part per trillion by volume) to several ppbv (parts per billion by volume) is Chemical Ionisation Mass Spectrometry (CIMS) using iodide ions to generate $I \cdot ClNO_2^-$ which can be detected at a mass-to-charge ratio (m/z) of 208 and 210 (McNeill et al., 2006). The first measurement highlighting the importance of $ClNO_2$ in the polluted marine boundary layer was performed by Osthoff et al. (2008) who detected mixing ratios exceeding 1 ppbv along the coast of Houston, Texas, originating from ship-plumes and urban and industrial NO_x sources. This was the starting point for numerous measurements of $ClNO_2$ in various locations around the globe with an initial focus on coastal areas in the United States (U.S.), e.g. the Los Angeles Basin in California (Riedel et al., 2012a; Wagner et al., 2012; Young et al., 2012). Other studies included coastal sites in Canada (Osthoff et al., 2018) and coastal / urban sites in the United Kingdom (Bannan et al., 2015; Bannan et al., 2017; Priestley et al., 2018; Sommariva et al., 2018). Whereas $ClNO_2$ was initially believed to play a significant role only in areas with marine influence (Behnke et al., 1997; Keene et al., 1999), mid-continental measurements in the U.S. (Thornton et al., 2010; Riedel et al., 2013; Faxon et al., 2015) revealed the importance of anthropogenic sources (e.g. industrial combustion, cooling towers, natural gas extraction and suspension of road salt) and sea salt chloride transported inland. Further studies reported significant mixing ratios of $ClNO_2$ at a semi-rural site in continental Germany (Phillips et al., 2012; Phillips et al., 2016) and at a mid-continental urban site in Canada (Mielke et al., 2011; Mielke et al., 2016). Observations at continental sites could be reproduced by a global model (Wang et al., 2019) when considering the transport of HCl (aq) which had been initially formed in the gas-phase through acid displacement in coastal regions. More recently, $ClNO_2$ at the > 1 ppbv level has been observed in the heavily industrialised North China Plain (Tham et al., 2016; Liu et al., 2017; Wang et al., 2017; Tham et al., 2018), with even larger mixing ratios measured in Beijing (Le Breton et al., 2018; Zhou et al., 2018) and Hong Kong (Wang et al., 2016).

The great variability seen in $ClNO_2$ mixing ratios in different locations reflects regional variability in its efficiency of production, which, as described above involves a complex set of chemical reactions, both in the gas- and particle phase and

which will vary over time and space. Most measurements of $ClNO_2$ to date have been measurements at single locations, though some data from mobile platforms such as aircraft (Mielke et al., 2013; Lee et al., 2018; McDuffie et al., 2018a; McDuffie et al., 2018b) and ships (Kercher et al., 2009; Riedel et al., 2012a) are available. With respect to understanding the formation and role of $ClNO_2$, much of the atmospheric boundary layer remains unexplored.

5 Here we present shipborne measurements of $ClNO_2$ in the marine boundary layer of the Mediterranean Sea and around the Arabian Peninsula, including the Red Sea and the Arabian Gulf. With a ship track from southern France to Kuwait we provide a unique marine $ClNO_2$ dataset with a large spatial coverage. This allows us to investigate the $ClNO_2$ production efficiency ε and its regional impact under various atmospheric conditions ranging from polluted marine and coastal environment to low-NO_x conditions in chemically aged air masses.

10 **2 Methods**

2.1 AQABA campaign

The measurements presented in this study were performed during the AQABA campaign which took place along the sea route between southern France and Kuwait in summer 2017. Five air-conditioned measurement containers with a variety of gas-phase and aerosol instrumentation were set up on-board the research vessel *Kommander Iona* which departed from

15 Southern France on 24 June 2017 and passed various regions including the Mediterranean Sea, the Suez Canal, the Red Sea, the Gulf of Aden, the Arabian Sea, the Gulf of Oman and the Arabian Gulf (see Fig. 2), reaching its destination Kuwait on 31 July 2017 (first leg) and covering a latitude / longitude span of 12–43 °N and 6–60 °E. After a short break in Kuwait the ship returned via the same route to southern France, arriving on 2 September 2017 (second leg). The trace-gases described in this paper were sampled from the centre of a common, high volume-flow inlet (10 m^3 min^{-1}, 0.15 m in diameter, 0.2 s

20 residence time) made of stainless steel, which was located on a measurement container at the front of the ship at a height of approximately 5.5 m above the foredeck.

Depending on the wind direction relative to the movement of the vessel, measurements were occasionally impacted by emissions from the stack of our own ship. Especially on the first leg, the relative wind direction was frequently from behind where the chimney was located. All datasets were filtered prior to analysis for periods where the measurements were

25 contaminated by stack emissions to avoid a potential bias in the results. The filter is based on short-term variation in NO and SO_2 signals and relative wind direction and reduces the useful data coverage to 58 % on the first leg and 95 % on the second leg.

2.2 Measurement of $ClNO_2$, HCl and SO_2

Nitryl chloride ($ClNO_2$), hydrogen chloride (HCl) and sulphur dioxide (SO_2) were detected with a Chemical Ionisation

30 Quadrupole Mass Spectrometer (CI-QMS) using an electrical, radio-frequency (RF) discharge ion-source. The instrument and the ion-molecule-reactions involved in the detection of the above-mentioned trace gases are described in detail by Eger

et al. (2019). Briefly, $ClNO_2$ was monitored as $I \cdot ClNO_2^-$ at a mass-to-charge ratio (m/z) of 208 and 210 subsequent to the reaction of $ClNO_2$ with I^- (McNeill et al., 2006; Osthoff et al., 2008; Thornton et al., 2010). $I \cdot ClNO_2^-$ is more specific than ICl^- (m/z 162 and 164) and has a lower background signal, providing a sensitivity of 0.61 Hz pptv^{-1} per 10^6 Hz of I^- at m/z 208 (and 0.20 Hz pptv^{-1} at m/z 210), a limit of detection (LOD) (2σ, 5 min) of 12 pptv and a total measurement uncertainty of 30 % ± 6 pptv. We chose the signal at m/z 208 for its higher signal-to-noise (S/N) ratio to calculate the $ClNO_2$ mixing ratios reported. For the whole campaign dataset, the ratio between m/z 208 and m/z 210 was 3.08 (R^2 = 0.96, see Fig. S1) which is very close to the expected value of 3.13 derived from the natural abundance of the ^{35}Cl and ^{37}Cl isotopes, indicating no significant interferences at either of the two m/z.

HCl was observed as $I(CN)Cl^-$ (m/z 188 and 190) (Eger et al., 2019) with a sensitivity of 0.17 Hz pptv^{-1} per 10^6 Hz of I^- at m/z 188 (and 0.05 Hz pptv^{-1} at m/z 190), a detection limit of 98 pptv and a total measurement uncertainty of 20 % ± 72 pptv. As m/z 190 suffers from known interferences (e.g. $I \cdot HNO_3^-$) and has a lower S/N ratio, we used m/z 188 to calculate the HCl mixing ratios reported. SO_2 was detected as ISO_3^- (m/z 207) with a sensitivity of 0.10 Hz pptv^{-1} per 10^6 Hz of I^-, a detection limit of 38 pptv and a total uncertainty of 20 % ± 23 pptv.

A flow of 2.5 slm (standard litres per minute) was drawn into the CI-QMS instrument via a ≈ 3 m long 6.35 mm (OD) PFA tubing while a 20 cm section of the inlet line in front of the IMR (ion molecule reactor) was heated to 200 °C to enable detection of peroxyacetyl nitrate (PAN) which is not reported here. ~~An additional bypass (1 slm) in front of the IMR pinhole was installed to improve response times.~~ The IMR region was held at a pressure of (18.00 ± 0.05) mbar by a dry vacuum scroll pump. The background signal was determined by periodically bypassing ambient air through a scrubber filled with steel wool where the trace gases of interest are efficiently destroyed at the hot surfaces (120 °C). To avoid condensation of water in the inlet lines in the containers, the pressure in the sampling line was reduced to ≈ 700–800 mbar with a bypass flow of ≈ 5 slm and by including an additional ≈ 50 cm long (coiled) piece of 3.18 mm (OD) PFA tubing. A 2 μm pore size membrane filter (Pall Teflo) was placed between high volume-flow inlet and CI-QMS sampling line to remove particles and was exchanged regularly to avoid accumulation of particulate matter. No indication for $ClNO_2$ formation via N_2O_5 reactions on salty surfaces in the inlet line was observed during AQABA, i.e. whenever we changed the particle filter or the inlet line, no change in signal was observed. Further the $ClNO_2$-to-N_2O_5 ratio was highly variable during AQABA (range of 0.35–59 with a median of 3.2) and $ClNO_2$ was occasionally measured in periods where no N_2O_5 was present.

$ClNO_2$ was calibrated twice during the campaign by simultaneously sampling a source of $ClNO_2$ via the CI-QMS and by a thermal dissociation cavity ring-down spectrometer (Sobanski et al., 2016). $ClNO_2$ was generated by passing Cl_2 over $NaNO_2$ as described previously (Thaler et al., 2011; Eger et al., 2019). HCl was calibrated four times throughout the campaign by adding a small flow over a permeation source to the main flow and monitoring the CI-QMS signal at m/z 188 and 190. SO_2 calibrations were performed seven times during the AQABA campaign by addition of a known flow of SO_2 from a gas cylinder (1 ppmv in synthetic air, Air Liquide). In contrast to $ClNO_2$ and HCl, correction of the SO_2 signal for its relative humidity (RH) dependence was necessary, which we derived from calibrations during AQABA where the RH was actively varied between 1 and 80 %.

The CI-QMS was operated in selected ion monitoring mode measuring mainly $ClNO_2$, HCl, SO_2, PAN and peracetic / acetic acid with a temporal resolution of approximately 15 s for each molecule. Changes in sensitivity were captured by permanently monitoring the primary ion signal (I^- and its water cluster) during ambient measurements and a background signal was recorded every 100 minutes. For further analysis, all data sets were averaged to 5 min temporal resolution. Our $ClNO_2$, HCl and SO_2 datasets provide about 12,500 data points distributed over 61.4 days with interruptions due to background determinations, calibrations, filter and gas bottle changes and instrument power-down at the harbours of Jeddah and Kuwait. For periods where the ship was in motion, the data coverage for all three trace gases was about 80 %.

2.3 Other trace gases

O_3 was measured by a commercial ozone monitor (2B Technologies, Model 202) based on optical absorption at 254 nm with a detection limit of 3 ppbv (10 s) and a total uncertainty of 2 % ± 1 ppbv. Mixing ratios of NO_x and NO_y ($NO_y = NO_x +$ reactive nitrogen trace gases + particulate nitrate) were monitored via Thermal Dissociation Cavity Ring-Down Spectroscopy (TD-CRDS) using a modified version of the instrument described by Thieser et al. (2016). The difference between the NO_y and the NO_x signal is referred to as NO_z, which includes organic nitrates (peroxyacetyl nitrates and alkyl nitrates), NO_3, N_2O_5, $ClNO_2$, HNO_3 and particulate nitrate. In contrast to Thieser et al. (2016), the TD-unit was operated at 850 °C to ensure detection of HNO_3 and nitrate in the particle phase. The detection limits for NO_x and NO_y were 80 and 160 pptv, respectively, with total uncertainties of 9 % ± 30 pptv. NO_z was calculated from measured NO_x and NO_y with a detection limit of 160 pptv and a total uncertainty of 13 % ± 42 pptv. NO_2 (LOD = 52 pptv (1s), total uncertainty = 7 %) and N_2O_5 (LOD = 6 pptv (1s), total uncertainty = 15 %) were measured by a five-channel TD-CRDS described by Sobanski et al. (2016). NO and NO_2 were measured by a ~~modified commercial~~ chemiluminescence detector (CLD 790 SR, ECO Physics, Duernten, Switzerland) (Fontijn et al., 1970; Li et al., 2015). The LOD (5 s) was 21 pptv for NO and 52 pptv for NO_2 and the total uncertainty 6 % respectively 7 %. The NO_2 data was in good agreement with the CRDS dataset ($R^2 = 0.95$) with a mean deviation of 6 %. The hydroxyl radical (OH) was measured using a Laser-Induced-Fluorescence method (Martinez et al., 2010; Novelli et al., 2014).

2.4 Meteorological parameter and actinic flux

Photolysis rates ($J_{O(1D)}$, J_{ClNO2} and J_{NO3}) were calculated from wavelength resolved actinic flux measured by a spectral radiometer (Metcon GmbH; Meusel et al. (2016)) located close to the common trace-gas inlet. Cross sections and quantum yields were taken from Burkholder et al. (2015). J-values were not corrected for upwelling UV radiation and are estimated to have an overall uncertainty of ≈ 10 %. A commercial NEPTUNE weather-station (Sterela) monitored various parameters such as temperature, relative humidity, wind speed and direction, speed of the vessel and GPS position.

2.5 Aerosols

An Aerosol Mass Spectrometer (Aerodyne HR-ToF-AMS, DeCarlo et al. (2006)) measured PM_1 non-refractory aerosol composition (30 s time resolution), including sulphate, nitrate, ammonium, chloride and total organics with an overall uncertainty of 35 %. An Optical Particle Spectrometer (OPC, Grimm model 1.109) measured the size distribution from 250 nm to 32 µm (6 s time resolution) with a total uncertainty of 25 %. A Fast Mobility Particle Spectrometer (FMPS, TSI model 3091) provided particle size distributions from 5.6 nm up to 560 nm (1 s time resolution). The particle surface area concentrations for PM_1 and PM_{10} were calculated from the OPC and FMPS datasets, the overall uncertainty of these variables is estimated to be 30 %. The inlet for the aerosol instrumentation was located at the top of a measurement container at a distance of $\approx$ 5 m to the common, trace-gas inlet described above. In order to avoid condensation in inlet lines, aerosol samples were passed through a drying system which reduced ambient relative humidity to an average value of $\approx$ 40 % in the measurement container. We calculated the ambient PM_1 particle surface area concentration (A) from the measured surface area concentration using a hygroscopic growth factor (on average 1.32 $\pm$ 0.24) based on ambient RH and aerosol composition. The calculation of the growth-factor is described in the supplementary information (Fig. S2–S5). The water soluble fraction of total suspended particles (TSP) was monitored with hourly resolution using a Monitor for AeRosols and Gases in Ambient Air, MARGA (Metrohm Applikon Model S2), sampling at a distance of $\approx$ 5 m from the common gas-phase inlet. In this work only results from Na^+ and Cl^- measurements (TSP), with detection limits equal to 0.05 and 0.01 µg m^{-3}, will be used.

3 Results and discussion

In the following, we use only data (5 min averages) which were free from contamination by the ship's own exhaust (see Sect. 2.1).

3.1 Overview of measurement regions and $ClNO_2$ mixing ratios observed

Figure 2 illustrates the ship's track during AQABA, divided into seven regions demarked by dashed black lines: The Mediterranean Sea, the Suez Canal including the Gulf of Suez, the Red Sea, the Gulf of Aden, the Arabian Sea, the Gulf of Oman and the Arabian Gulf. On the first leg, the CI-QMS measurements started south of Crete; on the second leg measurements terminated close to Sicily after two months of almost continuous measurement. Maximum $ClNO_2$ mixing ratios observed during each night ranged from the limit of detection to 586 pptv (see Fig. S6 for details). Figure 2 shows 1-hour averaged $ClNO_2$ mixing ratios along the ship track during (a) first and (b) second leg. Text boxes indicate the median night-time mixing ratios of O_3, HCl, NO_2 and SO_2 for the different regions where data from the first and second leg datasets have been combined. The night-time mean, median and range of the mixing ratios of these trace gases (and also of NO_2, temperature, relative humidity, NO_3 production rate and PM_1 particle surface area concentration) are listed in Table 1; a

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time-series of measured $ClNO_2$, HCl, SO_2 and O_3 is provided in Fig. S7 of the supplementary information. The predominant air-mass origin for each night was derived from 48 h back-trajectories calculated with HYSPLIT (Stein et al., 2015; Rolph et al., 2017) and is illustrated for both legs in Fig. S8 of the supplementary information. While Fig. 2 provides an overview of the measurements during both legs, Fig. 3a and 3b highlight 9-day periods indicating features that characterised the transition from one region to the next. Based on these three figures, we will discuss observed $ClNO_2$ mixing ratios and related parameters for the seven regions defined above.

Over the Mediterranean Sea, during periods when the CIMS was operational, we encountered mainly aged air masses which had passed over Italy, Greece or Turkey (Fig. S8) characterised by relatively high O_3 levels but low NO_2-to-NO_y ratios. As illustrated in Fig. 1 the formation of $ClNO_2$ is initiated by NO_3 production which will depend on O_3 levels and availability of NO_2. For the Mediterranean Sea the low NO_2 mixing ratios resulted in a weak NO_3 production term (Table 1) and low $ClNO_2$ mixing ratios. The only exceptions are two nights south of Sicily on the second leg where $ClNO_2$ mixing ratios up to 439 pptv were observed, which coincided with an increase in NO_2 originating from industrial sources on the mainland (Sicily and Italy). There are no previous measurements of $ClNO_2$ over the Mediterranean Sea but our data can be compared to the output of a regional model (Li et al., 2019) which predicts monthly average $ClNO_2$ mixing ratios up to 100 pptv in the south-eastern Mediterranean Sea and around Sicily, in broad agreement with our observations.

The Suez Canal and the Gulf of Suez were impacted by fresh emissions from ships, industry and urban centres with high NO_2, SO_2, HCl and $ClNO_2$ mixing ratios. On the night 22–23 August 2017 we measured the highest $ClNO_2$ mixing ratio of the whole campaign (586 pptv) due to exceptionally high NO_2 levels and NO_3 production rates.

Over the Red Sea, O_3 levels were elevated with the highest NO_3 production rates (up to 0.7 pptv s^{-1}) observed when approaching the Gulf of Suez with NO_x transported south from the region around the Suez Canal, the city of Cairo and the Sinai Peninsula (see back-trajectories in Fig. S8). $ClNO_2$ mixing ratios exceeded 200 pptv on most of the nights (with a maximum of 480 pptv) whereby elevated PM_1 particle surface area concentration and HCl mixing ratios indicated increased heterogeneous uptake and chloride availability.

Over the Gulf of Aden, with air mainly originating from Somalia, O_3 levels were close to 25 ppbv and NO_2 mixing ratios regularly exceeding 5 ppbv resulted in a NO_3 production rate up to $\approx$ 0.2 pptv s^{-1}. However, the PM_1 particle surface area concentration remained low and $ClNO_2$ was detected only occasionally (maximum of 379 pptv).

Over the Arabian Sea we experienced strong winds from the south with 48 h back-trajectories touching the coast of Somalia. $ClNO_2$ was generally below the detection limit (maximum 56 pptv) and NO_2, HCl and PM_1 particle surface area concentration were very low. Missing local sources of NO_x and low O_3 mixing ratios resulted in a weak NO_3 production term, partially responsible for the lack of $ClNO_2$. Low mixing ratios of $ClNO_2$ were occasionally detected that originated from single ships or point sources on the mainland.

Upon entering the Gulf of Oman, which marks the transition between remote marine environment and increased emissions from petrochemical industry and shipping lanes, NO_3 production rates increased significantly due to higher NO_x and O_3

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levels. $ClNO_2$ mixing ratios exceeded 200 pptv during two consecutive nights with a maximum value in this region of 376 pptv.

The Arabian Gulf was characterised by very high ozone levels (sometimes exceeding 150 ppbv) and SO_2 mixing ratios that generally exceeded 5 ppbv. For the first leg (sailing into the Arabian Gulf), the air-mass passed over Kuwait whereas for the second leg (sailing out of the Arabian Gulf) it mainly passed over Iran. In the Gulf region, which was heavily polluted by emissions from shipping and petrochemical industry, we also observed the highest HCl and PM_1 particle surface area concentration throughout the whole campaign. However, despite high NO_3 production rates ($\approx$ 0.4 pptv s^{-1}) due to NO_x emissions from oil and gas refineries as well as emissions from shipping and urban areas, we only observed relatively low $ClNO_2$ mixing ratios with a maximum value of 126 pptv close to Kuwait.

Consistent with Osthoff et al. (2008) we find significant amounts of nocturnal $ClNO_2$ in aged ship-plumes that could be identified by a defined peak-shape (Fig. S9) and covariance of NO_2 and SO_2 indicative of upwind point sources. As SO_2 and NO_2 are co-emitted from the combustion of ship fuel, it is not surprising that they show a co-variance. The consequence of the co-emission of NO_2 and SO_2 is that $ClNO_2$ is generally observed in the presence of both whereby high $ClNO_2$ mixing ratios were associated with aged ship-plumes.

Figure 4 shows diel profiles of nitryl chloride for the Red Sea and the Gulf of Oman together with the photolysis rate constant J_{ClNO2} illustrating that mixing ratios generally decreased at sunrise with a $ClNO_2$ lifetime of a few hours. Diel $ClNO_2$ profiles for other regions (see Fig. S10 of the supplementary information) look generally similar but with a varying maximum mixing ratio. Over the Red Sea $ClNO_2$ was often observed in plumes, whereas mixing ratios over the Gulf of Oman increased continuously after sunset indicating that we sampled a more homogeneously polluted air mass in which $ClNO_2$ accumulated over the course of the night. The median mixing ratio in the afternoon was still around 10 pptv, which we attribute to an HCl interference at m/z 208 and 210 described by Eger et al. (2019) rather than to the presence of $ClNO_2$ during the day when its production rate is close to zero and its lifetime is short due to photolysis. The magnitude of the HCl interference at the m/z used to monitor $ClNO_2$ was derived during HCl calibrations on-board the ship and found to be 0.006 Hz (pptv of HCl)$^{-1}$ which is about 1 % of the $ClNO_2$ count rate of 0.61 Hz (pptv of $ClNO_2$)$^{-1}$ at 10^6 Hz of I$^-$. However, during ambient air measurements the interfering signal was variable with a campaign average of (0.008 ± 0.005) Hz pptv^{-1}, which implies that a correction based on the HCl signal alone is not sufficient. The variable offset at the $ClNO_2$ mass contributes to the total measurement uncertainty and can be significant when analysing data close to the detection limit. Although on occasions several hundred pptv of $ClNO_2$ were observed, below we show that the $ClNO_2$ production efficiency was generally low. Reasons for this are examined in the following sections.

3.2 $ClNO_2$ yield per NO_3 molecule formed

We define the $ClNO_2$ production efficiency (ε) during AQABA as the number of $ClNO_2$ molecules generated per NO_3 molecule formed from the reaction of NO_2 with O_3 (Reaction R3). The instantaneous production rate of NO_3 is given by $k_1[NO_2][O_3]$ and the total number of NO_3 molecules formed over the course of the night is derived using a rate coefficient of

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$k_1 = 1.4 \times 10^{-13}$ exp (-2470/T) cm^3 molecule^{-1} s^{-1} (IUPAC, 2019) and integrating the NO$_3$ production term from the beginning of the night (t_0) to the time of the measurement (t) according to Eq. (2). In this calculation we assume that [NO$_2$] changes over time but [O$_3$] stays constant in good approximation.

$$[NO_3]_{int} = \int_{t_0}^{t} k_1 [O_3] \, [NO_2](t) \, dt \tag{2}$$

In order to account for the pre-sunset production of NO$_3$ at high solar zenith angles where N$_2$O$_5$ could already be detected, t_0 was defined as the point in time at which J_{NO3} was below 0.017 s^{-1} (about 10 % of maximum value during day). This was typically 30-50 minutes prior to sunset. All data points before sunset were however excluded from the analysis due to the increased uncertainty in the reaction time. The NO$_2$ mixing ratio at the beginning of the night, [NO$_2$]$_0$, was derived from the measured NO$_2$ mixing ratio at time t via Eq. (3) by assuming that NO$_2$ had been consumed by reaction with O$_3$ but the O$_3$ mixing ratio did not change significantly. Consequently, the amount of NO$_3$ produced along the air mass trajectory is equal to the difference between calculated [NO$_2$]$_0$ and measured [NO$_2$] (t).

$$[NO_2](t) = [NO_2]_0 \, e^{-k_1[O_3]t} \tag{3}$$

The ClNO$_2$ production efficiency ε can be determined by inserting the integrated NO$_3$ production over the course of the night and the measured ClNO$_2$ mixing ratio (assuming no losses) into Eq. (4).

$$\varepsilon = \frac{[ClNO_2]}{[NO_3]_{int}} \tag{4}$$

To account for fresh emissions of NO (e.g. by passing ships), the reaction time t' was calculated from Eq. (5) according to McDuffie et al. (2018a):

$$t' = (k_1[O_3]s) \, ln\left(\frac{[NO_y]}{[NO_2]}\right) \tag{5}$$

where s represents the number of NO$_2$ molecules required to make NO$_y$ and is 1 when NO$_3$ reacts directly with VOCs and 2 when NO$_3$ reacts with NO$_2$ to form N$_2$O$_5$, which subsequently hydrolyses to HNO$_3$. As discussed later, the direct NO$_3$ losses are dominant throughout the campaign compared to the heterogeneous N$_2$O$_5$ production, so to a good approximation, $s \approx 1$. As discussed by McDuffie et al. (2018a) inherent to the use of this expression is the assumption that NO$_y$ is conserved during the night; any losses of NO$_y$ (e.g. via deposition of HNO$_3$) lead to an underestimation of the true reaction time. As the calculated, night-time air mass age depends on the ratio between [NO$_2$] and [NO$_y$], the calculation breaks down whenever a fresh NO emission (e.g. from a nearby ship) is injected into an air-mass and unreacted NO is still present. To avoid this, we only analyse ClNO$_2$ data when NO is below the detection limit. In addition, we only consider data points where the calculated age of the air mass is equal to or exceeds the time elapsed since sunset as these air masses are unlikely to have been impacted by recent emissions. As the loss of NO$_z$ via deposition will result in an air mass age that is shorter than the true one, we relax the criterion for equality of reaction times by also including calculated air mass ages that are up to 25 % shorter (i.e. $t' \geq 0.75 \, (t - t_0)$). Assuming a deposition velocity of 1 cm s^{-1} (McDuffie et al., 2018a) for HNO$_3$ (which is the

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~~dominant component of NO$_\varepsilon$ during AQABA) and a boundary layer height of 1000 m results in a lifetime of HNO$_3$ with respect to deposition of about 28 h, i.e. within one night (typical duration of maximum 12 h) less than 35 % of the HNO$_3$ will be lost.~~ The reduced dataset provides 1742 data points with a median value of ε = 2.7 % for the whole campaign. The data reduction is described in more detail in the supplement (all the data shown in the manuscript corresponds to the application of method C), where the sensitivity of ε to these limitations and additional constraints is discussed. Here, we emphasise that even when limiting the dataset to ClNO$_2$ mixing ratios exceeding 100 pptv (see Table S1), the ClNO$_2$ production efficiency still remains relatively low (ε = 6.4 %, Fig. S11). Although the median values would be modified (Fig. S12), the relative differences between the regions and especially the low ε observed over the Arabian Gulf persist.

3.3 Temporal and regional variability in ε

To compare the efficiency of ClNO$_2$ formation in different regions, a median value for ε was derived for each individual night with the results illustrated in Fig. 5 in which the size of the circles (first leg) and stars (second leg) scales with the median NO$_3$ production rate (for a more detailed plot with 1-hour averaged data points see Fig. S13). Despite high NO$_3$ production rates (high O$_3$ and NO$_2$ levels), the lowest values of ε were observed over the Arabian Gulf, whereas elevated values of ε were found e.g. over the Arabian Sea, where NO$_3$ production was lowest.

Fig. 6 displays box-plots of ε for each region, calculated from between 41 and 546 data points per region. The median ClNO$_2$ production efficiency ε displays large night-to-night variability and interregional variability with the highest value found over the Gulf of Aden and the Arabian Sea (median = 4.7 %) and the lowest value found over the Arabian Gulf (median = 0.8 %). Median values of ε (in %) derived for the Mediterranean Sea, the Suez Canal, the Red Sea and the Gulf of Oman were 2.9, 2.7, 2.1 and 2.0. In the following, we examine the factors that cause the generally low efficiency in ClNO$_2$ production and also the regional variability in ε.

3.4 Factors influencing the ClNO$_2$ production efficiency

The uptake of N$_2$O$_5$ to aerosol particles can proceed via hydrolysis to HNO$_3$ as well as formation of ClNO$_2$, with yield f (see Fig. 1). Assuming no night-time losses, the concentration of ClNO$_2$ is given by Eq. (6). The overall ClNO$_2$ production efficiency ε as derived in Sect. 3.2 is dependent on f and on the relative rates of direct NO$_3$ loss (k_{dir}, Eq. 7) and indirect NO$_3$ loss (k_{het}, Eq. 8) where A is the particle surface area concentration, $\bar{c}$ is the mean molecular velocity of N$_2$O$_5$ ((24400 ± 160) cm s^{-1} during AQABA) and $K_{eq} = \frac{[\text{N}_2\text{O}_5]}{[\text{NO}_2]\,[\text{NO}_3]} = 2.8 \times 10^{-27}$ (T/300)$^{-0.6}$ exp(11000/T) cm^3 molecule^{-1} (IUPAC, 2019) is the temperature-dependent equilibrium constant (Reactions R4 and R5). During night-time, k_{dir} is determined by the NO$_3$ reactivity towards VOCs and NO, k_{het} by the rate of heterogeneous uptake of N$_2$O$_5$ to aerosol particles. In the following we calculate f and γ from our measurements, compare the values with the literature and quantify the contributions of k_{het} and k_{dir} to the overall NO$_3$ loss rate.

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$$[ClNO_2] = [NO_3]_{int} \times \varepsilon = [NO_3]_{int} \times \left(\frac{k_{het}}{k_{het}+k_{dir}}\right) \times f \tag{6}$$

$$k_{dir} = \sum_i (k_{VOC})_i [VOC]_i + k_{NO+NO3} [NO] \tag{7}$$

$$k_{het} = \gamma \left(\frac{A\bar{c}}{4}\right) K_{eq} [NO_2] \tag{8}$$

If no particulate chloride is available, f is zero and two NO_3^- ions are produced according to Reaction (R1) whereas if the particulate chloride concentration is large, f approaches unity and one NO_3^- anion plus one $ClNO_2$ molecule are formed. As particulate nitrate (NO_3^-) can leave the particle as HNO_3 we can write:

$$f = 2 \left(\frac{p[NO_3^-]+p[HNO_3]}{p[ClNO_2]} + 1\right)^{-1} \tag{9}$$

where p signifies a production rate. Equation (9) assumes that N_2O_5 uptake is the sole mechanism for the night-time production of HNO_3. In principal, as demonstrated by Phillips et al. (2016), f can be derived from field data using measured production rates (or concentrations) of inorganic nitrate (NO_3^- + HNO_3) and $ClNO_2$ according to Eq. (9). For AQABA, we derived the mixing ratio of total inorganic nitrate from measurements of NO_z (corrected for $ClNO_2$ and N_2O_5) assuming that, for this marine environment the contribution of organic nitrate in both gas- and particle phases is small compared to inorganic nitrate. f could be derived using Eq. (9) whenever there was a significant correlation (over a period of several hours) between $ClNO_2$ and inorganic nitrate, as illustrated in Fig. S14 for data obtained in the Gulf of Oman (25–26 July 2017) for which $f = 0.60 \pm 0.04$. The spatially and temporally variable sources of pollution during AQABA meant that requirement of a homogeneous fetch over periods of hours was rarely fulfilled and only a handful of values for f could be derived this way. Further values derived are 0.42 ± 0.06 (Red Sea, 15 July 2017), 0.84 ± 0.09 (Gulf of Oman, 24–25 July 2017) and 0.65 ± 0.05 (Mediterranean Sea, 29 August 2017.), indicating that the values of f were generally large whenever $ClNO_2$ was observed.

f can also be calculated (Eq. 10) if the relative concentrations of particulate chloride and water are known (Behnke et al., 1997; Bertram and Thornton, 2009):

$$f = \left(\frac{k_{Cl^-}([Cl^-])}{k_{Cl^-}[Cl^-] + k_{H2O}[H_2O]}\right) = \left(1 + \frac{[H_2O]}{\frac{k_{Cl^-}}{k_{H2O}}[Cl^-]}\right)^{-1} \tag{10}$$

where $\left(\frac{k_{Cl^-}}{k_{H2O}}\right) = 450$ is the ratio of rate constants (Ammann et al., 2013) for reaction of NO_2^+ (formed along with NO_3^- upon dissociation of N_2O_5 in the aqueous-phase) with either Cl^- or H_2O. The aerosol liquid water content $[H_2O]$ and chloride ion concentration $[Cl^-]$ were calculated using the E-AIM model (http://www.aim.env.uea.ac.uk/aim/model4/model4a.php) (Clegg et al., 1998; Friese and Ebel, 2010) using the ambient temperature and relative humidity along with PM_1 nitrate, sulphate, ammonium and chloride mass concentrations ($\mu g\ m^{-3}$) as reported by the AMS. From the median aerosol composition (PM_1) in the seven different regions we calculated median values of f, which are shown in Table 2. The values

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of f obtained via Eq. (10) were variable between regions, with medians of 0.53 in the Mediterranean Sea, 0.90 in the Suez Canal, 0.86 in the Red Sea, 0.76 in the Gulf of Aden, 0.87 in the Indian Ocean, 0.50 in the Gulf of Oman and 0.17 in the Arabian Gulf. To put these numbers in context, a value of $f = 0.9$ corresponds to a ≈ 1.1 mol L^{-1} Cl$^-$ solution. MARGA measurements of Cl$^-$ and Na$^+$ (total suspended particles, TSP) also indicated sea salt concentrations up to 20 μg m^{-3} in the

5   coarse mode over the Arabian Sea. A comparison of the [Cl$^-$] (TSP) from the MARGA (which detects NaCl as well as NH$_4$Cl) with the (largely) non-refractory [Cl$^-$] reported by the AMS reveals a strong co-variance. The AMS concentrations (PM$_1$) were on average $\approx 1\%$ of those reported by the MARGA (TSP) with the correlation between them indicating that the AMS chloride (PM$_1$) is mainly due to sea-salt rather than NH$_4$Cl. If we assume that $\approx 10\%$ of the sea salt mass (TSP) is associated with the fine mode (PM$_1$) as previously derived (Sommariva et al., 2018), we can use the MARGA (TSP) [Cl$^-$] to

10 estimate that the true PM$_1$ [Cl$^-$] would be about an order of magnitude higher than measured by the AMS. Under the assumption that this is true, f is > 0.67 in all seven regions implying that a lack of Cl$^-$ is not the reason for low values of ε, as may be expected for a marine environment.

Previous derivations of f in a marine environment (Texas coast) yield values between 0.1 and 0.65 (Osthoff et al., 2008) whereas even larger values (up to 0.9) have been reported for inland sites impacted by anthropogenic emissions (Young et

15 al., 2013) or by long-range transport of sea-salt ($0.035 < f < 1$) (Phillips et al., 2016). A median value of 0.138 ($0.003 < f < 1$) was derived for airborne measurements in a coastal region during winter (McDuffie et al., 2018b).

The uptake coefficient, γ, can be estimated using the parameterisation in Eq. (11) (Bertram and Thornton, 2009):

$$\gamma = B \times k \times \left(1 - \left(\left(a \times \frac{[H_2O(l)]}{[NO_3^-]} \right) + 1 + \left(b \times \frac{[Cl^-]}{[NO_3^-]} \right) \right)^{-1} \right) \tag{11}$$

where $B = 3.2 \times 10^{-8}$ s, $k = 1.15 \times 10^6 - 1.15 \times 10^6 \exp(-0.13\,[H_2O(l)])$ s^{-1} is the rate constant for the reaction N$_2$O$_5$(aq) +

20 H$_2$O(l), $a = 0.06$ denotes the ratio of rate constants for reactions H$_2$NO$_3^+$(aq) + H$_2$O(l) and H$_2$NO$_3^+$(aq) + NO$_3^-$(aq) and $b = 29$ denotes the ratio of rate constants for reactions H$_2$NO$_3^+$(aq) + Cl$^-$ and H$_2$NO$_3^+$(aq) + NO$_3^-$(aq). Using AMS data for [NO$_3^-$] and [Cl$^-$] we derive $\gamma = 0.033 \pm 0.003$ where the standard deviation encompasses the weak inter-regional variation (see Table 2). This value is consistent with $\gamma = 0.03 \pm 0.02$ reported by Aldener et al. (2006) for a polluted marine environment. The large values of γ reflect the low PM$_1$ particulate nitrate concentrations observed during AQABA where the high temperatures

25 favour the partitioning of particulate nitrate and HNO$_3$(g) to the gas-phase. A suppression of γ through the presence of organics in the particle phase has been reported (Bertram et al., 2009), though the low (generally < 2) organic-to-sulphate ratio in particles observed during AQABA suggest that this is not likely to be important for the present analysis.

The campaign averaged, fractional contribution of coarse mode particles (PM$_{10}$ - PM$_1$) to the total particle surface area concentration (A) was only (14 $\pm$ 14) % so that the uncertainty incurred when using the PM$_1$ particle surface area

30 concentration to derive the heterogeneous NO$_3$ loss rate (k_{het}, Table 2) from Eq. (8) is negligible. However, during two

periods (of 2–5 days duration) when the ship was sailing through the Gulf of Aden / Arabian Sea / Southern Red Sea and air masses originated from the deserts of Eritrea, Djibouti and Ethiopia, the contribution from the coarse mode particles to the aerosol surface area increased to about 60 % mainly due to dust (and sea-salt). The relative contribution of dust and sea-salt to the coarse mode was estimated using MARGA measurements of Ca^{2+} and Na^+ in TSP. The dust loading was calculated by

5 assuming that the dust aerosols of Saharan origin consist of 10 % calcium (Molinaroli et al., 1993). Freshly generated dust particles do not contain chloride and the uptake of N_2O_5 to them ($\gamma = 0.02 \pm 0.01$ for Saharan dust, Tang et al. (2012)) does not result in $ClNO_2$ formation ($f = 0$) but contributes to k_{het} through the additional aerosol surface area, thus lowering the $ClNO_2$ production efficiency. In contrast, the uptake of N_2O_5 to coarse mode sea-salt particles has a $ClNO_2$ yield f close to unity (Ammann et al., 2013), and can therefore enhance formation of $ClNO_2$. Throughout the whole campaign, the

10 contribution of dust to the uptake of N_2O_5 to coarse mode particles was much larger than that of sea-salt (on average (13 ± 10) % with a range of 0–40 %). A close examination of the periods (14–19 July and 15–17 August 2017) strongly influenced by dust particles did not reveal any (anti)correlation between elevated dust concentration and $ClNO_2$ mixing ratios. This is also reflected by the relatively high median value of ε over the Gulf of Aden where the highest concentrations of coarse mode particles were observed.

15 Using a campaign average of $\varepsilon = 2.7$ % and a maximum value of $f = 1$ we can show that, with 97.3 %, the direct (gas-phase) loss of NO_3 (k_{dir}) is much more important than indirect losses via N_2O_5 uptake (k_{het}), which contribute the remaining 2.7 %. Assuming a very conservative estimate of $f = 0.5$ would still result in a contribution of 94.6 respectively 5.4 %. To put this into context, over the Gulf of Oman we calculated a median value of $k_{het} = 6.5 \times 10^{-4}$ s^{-1} (Table 2) which would result (assuming $\varepsilon = 2.0$ % and $f = 0.5$) in a direct NO_3 loss term towards VOCs of $k_{dir} = 1.6 \times 10^{-2}$ s^{-1}. The relatively small

20 contribution of the heterogeneous loss term is readily explained by high mean night-time temperatures of 25–35 °C during AQABA, which favour the existence of NO_3 rather than N_2O_5. To illustrate this, we calculate that a nocturnal temperature of 20 °C (rather than 30 °C) would increase the contribution of k_{het} to 21% of the total NO_3 loss and lead to 3 times higher $ClNO_2$ mixing ratios assuming that the rate constants for reaction between NO_3 and reactive trace gases are not strongly temperature-dependent. A further reduction in temperature to 10 °C would lead to equality in k_{het} and k_{dir} and result in a

25 factor of seven more $ClNO_2$ than observed.

The direct (gas-phase) loss rate of NO_3 can also be calculated from Eq. (7) if the concentrations of all VOCs contributing to its reactivity are known. However, for a large fraction of each night NO_3 was below the detection limit (ca. 5 pptv) despite a high production rate (large mixing ratios of NO_2 and O_3). A steady-state analysis of NO_3 production and loss indicated a high total reactivity which could not be attributed to measured trace gases (k_{dir}) or heterogeneous losses of N_2O_5 (k_{het}). A detailed

30 analysis of the NO_3 lifetime and the role of VOCs is beyond the scope of the present manuscript and will be described in detail in a separate publication.

3.5 Comparison with literature

In the following, we compare the generally low values of ε derived during AQABA with previous determinations. Mielke et al. (2013) report a yield of $ClNO_2$ relative to the total amount of NO_3 formed at night of 0.7 % to 62 % with a median of 12% in the polluted coastal boundary layer in Pasadena, California. In contrast, $ClNO_2$ production efficiencies derived for the urban boundary layer of Calgary, Canada (Mielke et al., 2016), were significantly lower, ranging from 0.1 % to 4.5 % (10th and 90th percentiles, median 1.0 %). Osthoff et al. (2018) report very low efficiencies with a median of 0.17 % and a maximum of 5.4 % for the Lower Fraser Valley of British Columbia, Canada, potentially due to a lack of available aerosol chloride. For AQABA we derive a region-dependent median efficiency of 1–5 % with a campaign median of 2.8 %, despite similar conditions to Mielke et al. (2013), i.e. mostly polluted marine environment. The difference can be attributed to exceptionally high nocturnal temperatures during AQABA with a median of 25–35 °C for different regions, shifting the equilibrium from N_2O_5 towards NO_3 and favouring direct NO_3 losses. For comparison, daily minimum temperatures during the study reported by Mielke et al. (2013) were 10–20 °C (Ryerson et al., 2013) whereas the lowest temperature measured during the whole AQABA cruise was 22 °C. Based on the $ClNO_2$ dataset reported by Phillips et al. (2012) for continental Germany, where $ClNO_2$ mixing ratios up to 800 pptv were reported, we calculate values of ε that range from 0.4 % to 12.3 % (10th and 90th percentiles) with a median of 2.6 %. On nights where $ClNO_2$ mixing ratios above 100 pptv were observed, range and median increase to 5.0–24.1 % and 10.6 % respectively. Compared to AQABA the yield per NO_3 molecule formed on nights where $ClNO_2$ was present at levels > 100 pptv is about a factor of 2 higher for this dataset, again most likely a result of the lower nocturnal temperatures.

3.6 Cl atom generation from $ClNO_2$ and HCl

In this section we assess the role of two gas-phase chlorine reservoirs, $ClNO_2$ and HCl as sources of Cl atoms during AQABA. Other potential Cl sources (e.g. Cl_2 photolysis) are not considered here as we do not have experimental data to quantify their impact. Compared to the complex route to $ClNO_2$ formation described above (Fig. 1), the formation of HCl in the polluted marine environment can be traced back to its displacement from sea-salt particles by stronger acids, such as H_2SO_4 and HNO_3 (Keene et al., 1999). The high emission rates of NO_x and SO_2 by ship traffic resulted in enhanced concentrations of both NO_2 and SO_2 (see Fig. S9) during parts of the AQABA campaign. Both NO_2 and SO_2 are oxidised via OH to form HNO_3 and H_2SO_4, both of which can be taken up by sea-salt containing aerosol releasing HCl. The release of HCl through acid displacement leads to a deficit in particulate Cl^- concentrations which can be expressed in terms of a chloride depletion factor (Eq. 12) where $[Na^+]$ and $[Cl^-]$ represent the concentrations in mol m^{-3} and 1.174 is the molar ratio of Cl^- to Na^+ found in sea-water (Zhuang et al., 1999).

$$\text{Cl depletion [\%]} = 100 \times \frac{(1.174 \times [Na^+] - [Cl^-])}{1.174 \times [Na^+]} \tag{12}$$

In Fig. 7 we present a time series (Gulf of Oman / Arabian Gulf) in which significant differences in particulate Na^+ and Cl^- concentrations coincide with high mixing ratios of NO_2, SO_2 and HCl (chloride depletion up to 90 %) indicating efficient HCl acid displacement by HNO_3 and H_2SO_4 in this region.

The instantaneous production rate of Cl atoms from the photolysis of $ClNO_2$ (pCl_{ClNO2}) is given by the photolysis rate constant for $ClNO_2$ (J_{ClNO2}) and its concentration (Eq. 13) whereas the instantaneous Cl production rate from HCl (pCl_{HCl}) requires knowledge of the OH concentration (Eq. 14) and the rate coefficient for reaction between OH and HCl (k_{OH+HCl} = 1.7×10^{-12} exp(-230/T) cm^3 molecule^{-1} s^{-1} (Atkinson et al., 2007; IUPAC, 2019)).

$$pCl_{ClNO2} = J_{ClNO2} [ClNO_2] \tag{13}$$

$$pCl_{HCl} = k_{OH+HCl} [OH][HCl] \tag{14}$$

In the following analysis, we focus on two consecutive nights in the Gulf of Oman region (Fig. 8) where we observed a monotonous increase of $ClNO_2$ mixing ratios up to $\approx$ 300 pptv during the second half of the night followed by a decrease over a 4 hour period starting at sunrise (upper panel). The corresponding Cl production rate from $ClNO_2$ photolysis reaches a maximum of 0.8×10^6 molecule cm^{-3} s^{-1} on the first night and 0.7×10^6 molecule cm^{-3} s^{-1} on the second night.

To place this in context, we also make a rough estimate (lower limit) to the rate of OH radical production (pOH_{O3}) from the photolysis of O_3 in the presence of H_2O (Eq. 15).

$$pOH_{O3} = \frac{2 J_{O1D} [O_3] \times k_{H2O} [H_2O]}{k_{H2O} [H_2O] + k_{N2} [N_2] + k_{O2} [O_2]} \tag{15}$$

where J_{O1D} is the photolysis rate constant for O_3, k_{H2O} = 2.4×10^{-11} cm^3 molecule^{-1} s^{-1}, k_{N2} = 2.15×10^{-11} exp(110/T) cm^3 molecule^{-1} s^{-1} and k_{O2} = 3.2×10^{-11} exp(67/T) cm^3 molecule^{-1} s^{-1} (IUPAC, 2019) refer to reactions of $O(^1D)$ with H_2O, N_2 and O_2, respectively. As we do not consider other OH production channels (e.g. photolysis of HONO or HO_2 + NO), which can be of importance under more polluted conditions, pOH_{O3} represents a lower limit of pOH. Although the maximum, midday OH production rates from O_3 photolysis ($\approx 1 \times 10^7$ OH molecule cm^{-3} s^{-1}) are about an order of magnitude higher than Cl atom production rates, during the first $\approx$ 2 hours after sunrise, pCl_{ClNO2} and pOH_{O3} are roughly equal for this particular case-study.

In order to examine the regional dependence of Cl formation we calculated the $ClNO_2$ production rate from the regional median values of ε, the NO_3 production rate and the length of the night as described previously. As $ClNO_2$ is completely photolysed to Cl during daytime, the total Cl-production from $ClNO_2$ photolysis (ΣCl_{ClNO2}) is equal to the median amount of $ClNO_2$ formed during the night. The total Cl-production from HCl + OH (ΣCl_{HCl}) was calculated by integrating the production rate pCl_{HCl} (based on [HCl] and [OH] measurements, Eq. 14) over the median diurnal profile. To calculate the contribution of O_3 photolysis to the OH production (ΣOH_{O3}), we integrated pOH_{O3} (calculated via Eq. 15) over the median diel profile. In Table 3 we summarise the average, total daytime production of OH and Cl in the seven regions. The median radical production over one diurnal cycle calculated for the Suez region is exemplified in Fig. 9, whereby we assume an

initial 363 pptv of $ClNO_2$ at sunrise (Table 3), which is subsequently photolysed according to J_{ClNO2}. Considering the entire campaign, we conclude that Cl-atom formation via $ClNO_2$ photolysis and OH-initiated HCl oxidation are of similar magnitude, though their relative contributions show large regional variability. For example, Cl formation from OH + HCl (ΣCl_{HCl}) is a factor ≈ 10 more important than $ClNO_2$ photolysis (ΣCl_{ClNO2}) over the Mediterranean Sea where the NO_3 (and

5　thus $ClNO_2$) production rate was low owing to low NO_x levels and where the OH concentrations were highest. This is consistent with Li et al. (2019) who indicate the importance of HCl over the eastern Mediterranean Sea for which monthly average mixing ratios of 0.5–1.5 ppbv were predicted by a regional model. ΣCl_{HCl} is also a factor ≈ 2 higher than ΣCl_{ClNO2} over the Arabian Gulf where the $ClNO_2$ production efficiency, ε, was low. In all other regions ΣCl_{ClNO2} was about equal to ΣCl_{HCl} or higher by factors between ≈ 1 and 4. This in in broad agreement with Riedel et al. (2012a) who report roughly

10　equal importance of $ClNO_2$ and HCl as chlorine atom sources in the polluted marine boundary layer of the Los Angeles region. They also report a maximum midday Cl production rate from $ClNO_2$ photolysis of 0.6 x 10^6 molecule cm^{-3} s^{-1}, which is similar to the production rates we obtained over the Gulf of Oman (Fig. 8).

When comparing the total number of chlorine atoms generated during the day ($\Sigma Cl_{total} = \Sigma Cl_{ClNO2} + \Sigma Cl_{HCl}$) with the total number of OH formed from O_3 photolysis (ΣOH_{O3}) over the same period, we find the expected domination of OH for all

15　regions. The largest contribution of Cl to the total radical production (4 %) was observed over the Suez Canal where the Cl production was highest (about 32 % from HCl and 68 % from $ClNO_2$). The lowest ratio of Cl-to-OH production was observed over the Arabian Sea, reflecting the low levels of NO_x and very low rates of NO_3 generation. Although only (at maximum) ≈ 4 % of the radicals generated are Cl atoms, they react more rapidly than OH with some classes of hydrocarbons, especially saturated hydrocarbons and small oxygenates. For example, the relative, room temperature rate

20　coefficients (k_{Cl} / k_{OH}) are 16, 61, 127 and 242 for reaction with CH_4, CH_3OH, C_3H_8 and C_2H_6, respectively (Atkinson et al., 2004; IUPAC, 2019). The impact of chlorine atoms is thus mainly seen in the oxidation rates of unsaturated hydrocarbons, the relative concentrations of which may be modified according to the relative abundance of OH and Cl and the relative reaction rate constants. For the AQABA campaign, evidence for such effects has been reported by Bourtsoukidis et al. (2019), and for the global scale Wang et al. (2019) conclude that oxidation by Cl atoms accounts for 1.0 % of methane loss

25　with larger impacts on ethane (20 %), propane (14 %), and methanol (4 %).

4 Summary and conclusion

The AQABA campaign, which took place in summer 2017 along the sea route from southern France to Kuwait provided the first $ClNO_2$ measurements in the marine boundary layer of the Mediterranean Sea, the Suez Canal, the Red Sea, the Gulf of Aden, the Arabian Sea, the Gulf of Oman and the Arabian Gulf. Along the ship track we observed a large variability in

30　$ClNO_2$ mixing ratios with nocturnal maxima ranging from below the detection limit over the Indian Ocean to a few hundred pptv over the Gulf of Oman, the northern part of the Red Sea, the Gulf of Suez and the Mediterranean Sea close to Sicily / Italy with a campaign maximum of ≈ 600 pptv observed over the Gulf of Suez.

The overall ClNO$_2$ production efficiency, i.e. the yield of ClNO$_2$ per NO$_3$ molecule formed in the reaction of NO$_2$ with O$_3$, was generally low (median of 2.7 % for the whole campaign) and highly variable within individual nights and between different regions with values (in percent) of 2.9, 2.7, 2.1, 4.7, 4.7, 2.0, and 0.8 over the Mediterranean Sea, the Suez Canal, the Red Sea, the Gulf of Aden, the Arabian Sea, the Gulf of Oman and the Arabian Gulf, respectively. The relatively low

5   ClNO$_2$ production efficiency compared to previous measurements in the polluted marine boundary layer or at continental sites was attributed to high nocturnal temperatures during AQABA (25–35 °C), which significantly shifted the equilibrium between NO$_3$ and N$_2$O$_5$ towards NO$_3$ and lowered the importance of N$_2$O$_5$ uptake to particles relative to direct NO$_3$ losses. The low ClNO$_2$ production efficiency in the Arabian Gulf (< 1 %) results from a combination of high temperatures, enhanced NO$_3$ reactivity and lowered chloride availability. The photolysis of ClNO$_2$ was found to represent an important

10   source of chlorine radicals in the early morning in areas where efficient night-time production was observed, and was augmented (and sometimes exceeded) by Cl atoms formation from the reaction of OH with HCl, especially in areas where ppbv levels of HCl were observed such as the Mediterranean Sea or the Arabian Gulf. Although the amount of Cl atoms generated were found to be a factor 25 to 300 less than the amount of OH molecules generated from O$_3$ photolysis, the high rate coefficients ratio for Cl compared to OH reactions towards some hydrocarbons imply that Cl may enhance hydrocarbon

15   oxidation, especially in the early morning.

Data availability

The data are available upon request to all scientists agreeing to the AQABA protocol from August 2019 (https://doi.org/10.5281/zenodo.3369035).

Author contributions

20   Philipp Eger performed the CI-QMS measurements of ClNO$_2$, HCl and SO$_2$ during the AQABA campaign, evaluated the field data and wrote the manuscript. John Crowley operated the CI-QMS and the CRDS-instruments during parts of the first leg and supervised the study. NO$_2$ and N$_2$O$_5$ data were provided by Justin Shenolikar. NO$_x$, NO$_y$, and NO$_z$ data were provided by Nils Friedrich. J-values were measured by Jan Schuladen. Ivan Tadic and Horst Fischer contributed the NO and NO$_2$ datasets. AMS, OPC and FMPS measurements and analysis were performed by James Brooks, Eoghan Darbyshire,

25   Friederike Fachinger and Frank Drewnick. MARGA data was provided by Michael Pikridas and Jean Sciare. Roland Rohloff, Sebastian Tauer, Monica Martinez and Hartwig Harder provided the OH dataset. Jos Lelieveld designed the AQABA campaign. All authors contributed to the manuscript.

Competing interests

The authors declare that they have no conflict of interest.

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Table 1: Measured trace gases and other parameters (5 min, night-time only) for the different regions indicated in Fig. 2.

Parameter[1]	Med. Sea[2]	Suez	Red Sea	Aden	Arab. Sea	Oman	Arab. Gulf
ClNO$_2$ [pptv]	N = 751	N = 376	N = 1247	N = 451	N = 918	N = 475	N = 500
Mean ± STD[1]	20 ± 39	75 ± 101	47 ± 55	41 ± 56	7 ± 8	67 ± 71	21 ± 18
median	9	36	22	19	6	39	18
range	LOD–439	LOD–586	LOD–480	LOD–379	LOD–56	LOD–376	LOD–126
HCl [ppbv]	N = 749	N = 375	N = 1245	N = 449	N = 918	N = 473	N = 498
Mean ± STD	0.68 ± 0.52	0.94 ± 0.63	0.81 ± 0.36	0.28 ± 0.46	0.00 ± 0.13	0.86 ± 0.70	1.20 ± 0.99
median	0.71	0.98	0.76	0.12	LOD	0.71	0.92
range	LOD–3.09	LOD–3.54	LOD–2.58	LOD–1.98	LOD–0.38	LOD–3.23	LOD–4.49
SO$_2$ [ppbv]	N = 751	N = 376	N = 1247	N = 451	N = 918	N = 475	N = 500
Mean ± STD	0.64 ± 0.65	2.84 ± 3.87	0.81 ± 0.77	1.26 ± 2.89	0.18 ± 0.65	2.38 ± 2.26	3.71 ± 2.01
median	0.46	1.57	0.66	0.55	0.07	1.39	3.25
range	0.06–7.24	0.06–34.98	0.04–9.75	LOD–30.34	LOD–12.33	0.12–12.34	0.80–14.85
O$_3$ [ppbv]	N = 751	N = 376	N = 1247	N = 451	N = 800	N = 475	N = 500
Mean ± STD	61.6 ± 10.2	45.4 ± 16.1	58.1 ± 13.8	28.0 ± 10.9	24.4 ± 3.4	36.0 ± 16.4	83.9 ± 36.2
median	63.4	46.3	59.6	25.7	25.6	30.0	69.8
range	37.8–82.8	1.7–88.1	10.6–90.3	3.1–55.3	1.3–29.6	10.2–83.0	24.2–163.3
NO$_2$ [ppbv]	N = 547	N = 288	N = 1217	N = 449	N = 898	N = 470	N = 500
Mean ± STD	0.80 ± 1.57	7.37 ± 7.56	1.97 ± 3.50	2.94 ± 4.11	0.63 ± 1.57	4.57 ± 3.93	1.59 ± 2.34
median	0.21	5.27	0.97	1.49	0.21	3.59	0.80
range	0.01–14.45	0.24–44.03	0.01–56.81	0.07–25.87	0.01–20.88	0.12–30.38	0.16–23.22
*p*NO$_3$ [pptv s^{-1}]	N = 547	N = 288	N = 1138	N = 449	N = 783	N = 470	N = 500
Mean ± STD	0.04 ± 0.07	0.27 ± 0.21	0.10 ± 0.13	0.06 ± 0.05	0.01 ± 0.02	0.16 ± 0.14	0.12 ± 0.13
median	0.01	0.23	0.05	0.04	0.004	0.13	0.08
range	0.00–0.82	0.01–1.01	0.00–1.23	0.00–0.25	0.00–0.18	0.00–0.91	0.01–1.19
T [°C]	N = 751	N = 376	N = 1247	N = 451	N = 918	N = 475	N = 500
Mean ± STD	26.7 ± 0.6	27.2 ± 1.7	31.3 ± 1.7	31.3 ± 1.1	25.4 ± 0.9	32.7 ± 1.3	34.4 ± 0.6
median	26.8	27.5	31.8	31.1	25.5	33.0	34.4
range	25.5–29.1	21.6–31.7	27.7–34.1	29.3–35.6	23.5–31.2	29.6–35.8	33.2–36.4
RH [%]	N = 751	N = 376	N = 1247	N = 451	N = 918	N = 475	N = 500
Mean ± STD	79.0 ± 5.6	72.4 ± 9.1	73.2 ± 6.3	74.6 ± 10.2	88.8 ± 2.9	79.8 ± 8.9	76.5 ± 7.9
median	79.5	71.6	72.7	75.6	89.1	81.5	77.8
range	63.2–89.3	40.6–93.2	58.7–88.4	27.2–90.0	79.3–94.7	51.9–95.1	53.3–91.6
A [μm^2 cm^{-3}]	N = 690	N = 373	N = 1101	N = 440	N = 899	N = 222	N = 393
Mean ± STD	202 ± 61	240 ± 200	243 ± 62	111 ± 64	81 ± 62	363 ± 317	1010 ± 600
median	178	166	257	107	77	275	809
range	114–466	90–1134	111–376	35–338	23–426	101–1244	179–2726

Notes: [1] Parameter: STD = standard deviation; RH = relative humidity; *A* = ambient PM$_1$ particle surface area concentration (see section 2.5); *p* denotes a production rate; *N* = number of data points; LOD = Limit of detection (see Sect. 2.2). [2] Regions: Mediterranean Sea (Med. Sea): 30 June–01 July and 24–30 August. Suez Canal and Gulf of Suez (Suez): 02–03 July and 22–23 August. Red Sea: 03–15 July and 17–21 August. Gulf of Aden (Aden): 16–18 July and 15–16 August. Arabian Sea (Arab. Sea): 19–23 July and 07–14 August. Gulf of Oman (Oman): 24–27 July and 05–06 August. Arabian Gulf (Arab. Gulf): 28 July–04 August 2017.

Table 2: Regional variability in $ClNO_2$ production efficiency (ε), $ClNO_2$ yield (f), N_2O_5 uptake coefficient (γ) and heterogeneous NO_3 loss rate (k_{het}).

Region	ε [%]	f [a]	γ [b]	k_{het} [10^{-5} s^{-1}]
Med. Sea	2.9	0.53	0.034	4.4
Suez	2.7	0.90	0.031	69.7
Red Sea	2.1	0.86	0.031	12.9
Gulf of Aden	4.7	0.76	0.031	8.5
Arab. Sea	4.7	0.87	0.036	1.8
Gulf of Oman	2.0	0.50	0.033	64.8
Arab. Gulf	0.8	0.17	0.036	34.3

Notes: [a]Calculated from Eq. (10); [b]Calculated from Eq. (11).

5

Table 3: Regional variability in ε, pNO_3 and OH and Cl radical production integrated over one diel cycle.

Region	ε [%]	pNO_3 [pptv s^{-1}]	ΣCl_{ClNO2} [pptv]	ΣCl_{HCl} [pptv]	ΣOH_{O3} [pptv]	[1]$\Sigma Cl_{total}/\Sigma OH_{O3}$ [%]
Med. Sea	2.9	0.012	18	234	12 364	2.0
Suez	2.7	0.231	363	170	13 216	4.0
Red Sea	2.1	0.053	48	53	12 411	0.8
Gulf of Aden	4.7	0.043	96	24	5 608	2.1
Arab. Sea	4.7	0.004	11	3	4 639	0.3
Gulf of Oman	2.0	0.130	155	127	12 649	2.2
Arab. Gulf	0.8	0.077	25	50	10 985	0.7

Notes: ΣCl_{ClNO2} and ΣCl_{HCl} are the integrated formation of Cl atoms from $ClNO_2$ and HCl respectively. ΣOH_{O3} is the integrated formation of OH from O_3 photolysis. [1]$\Sigma Cl_{total} = \Sigma Cl_{ClNO2} + \Sigma Cl_{HCl}$.

$$NO_2 \xrightarrow{O_3} NO_3 \underset{}{\overset{NO_2}{\rightleftarrows}} N_2O_5 \xrightarrow{\gamma \cdot f} ClNO_2$$

$$\text{VOCs} \downarrow \qquad\qquad \downarrow \gamma \cdot (1-f)$$

$$NO_3^-$$

with ε over the top from NO_3 to $ClNO_2$.
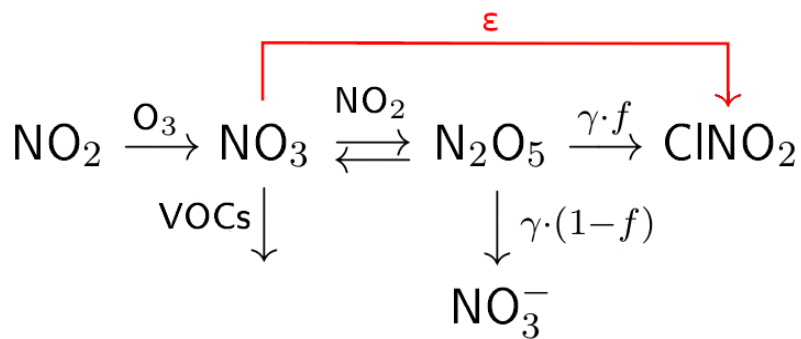
Figure 1: Simplified scheme of chemical reactions and parameters involved in the formation of $ClNO_2$. The $ClNO_2$ production efficiency per NO_3 formed is denoted as ε. The uptake coefficient is represented by γ and the $ClNO_2$ yield (per N_2O_5 taken up) by f.
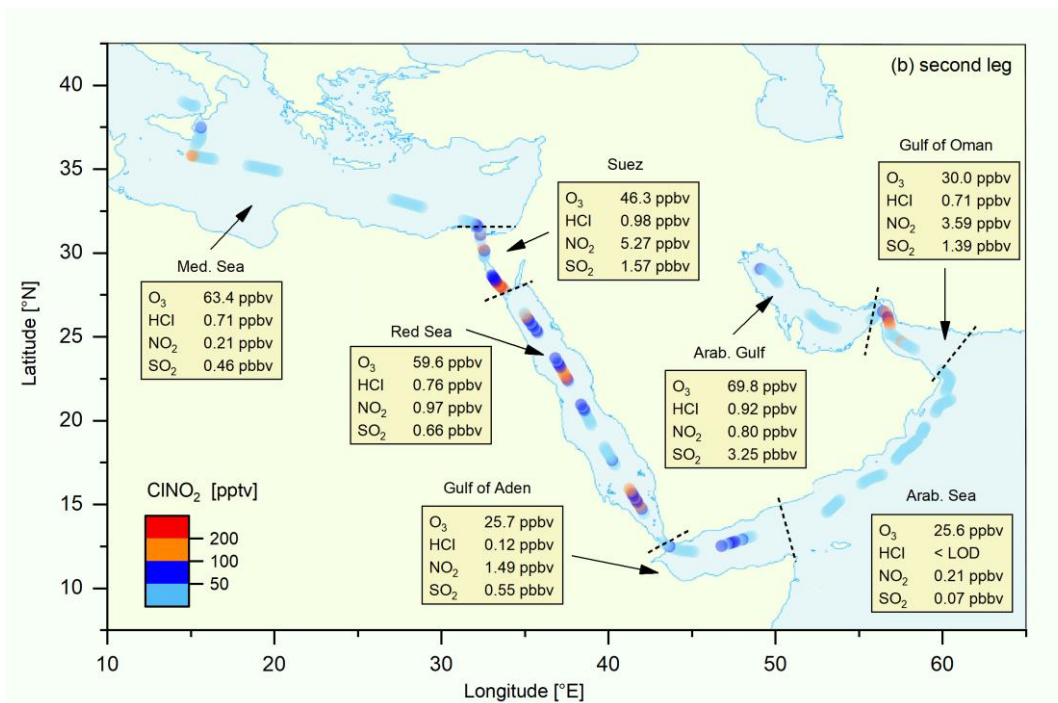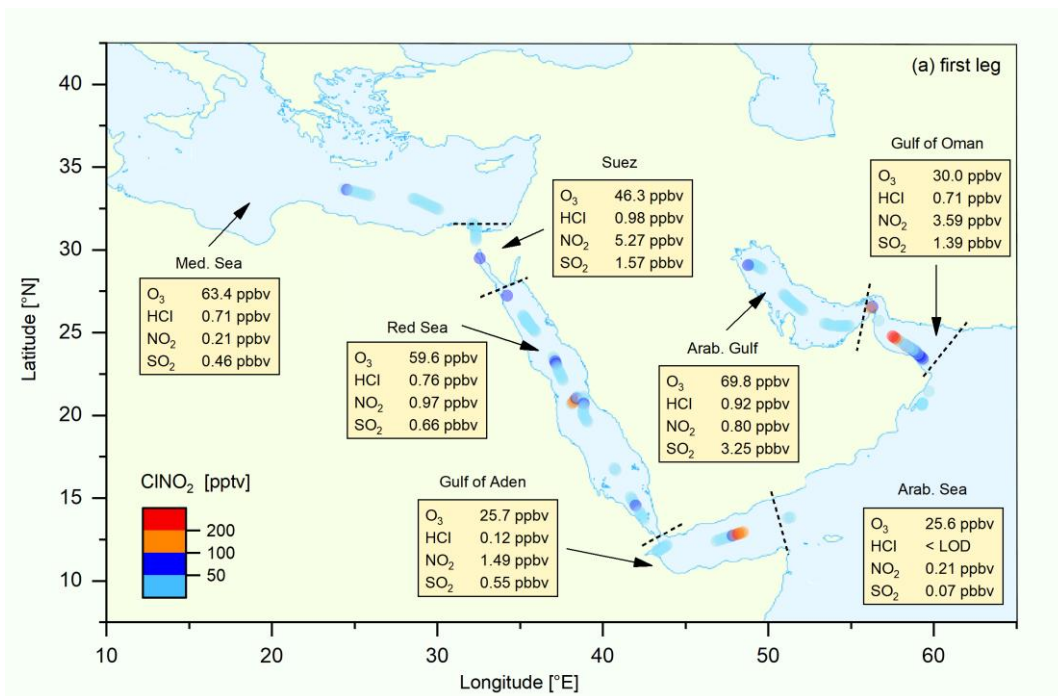
Figure 2: Map of nocturnal ClNO$_2$ mixing ratios (1-hour averages) for (a) first and (b) second leg, together with (median) night-time mixing ratios of O$_3$, HCl, NO$_2$ and SO$_2$ for different regions (data from both legs combined) demarked by the dashed lines.
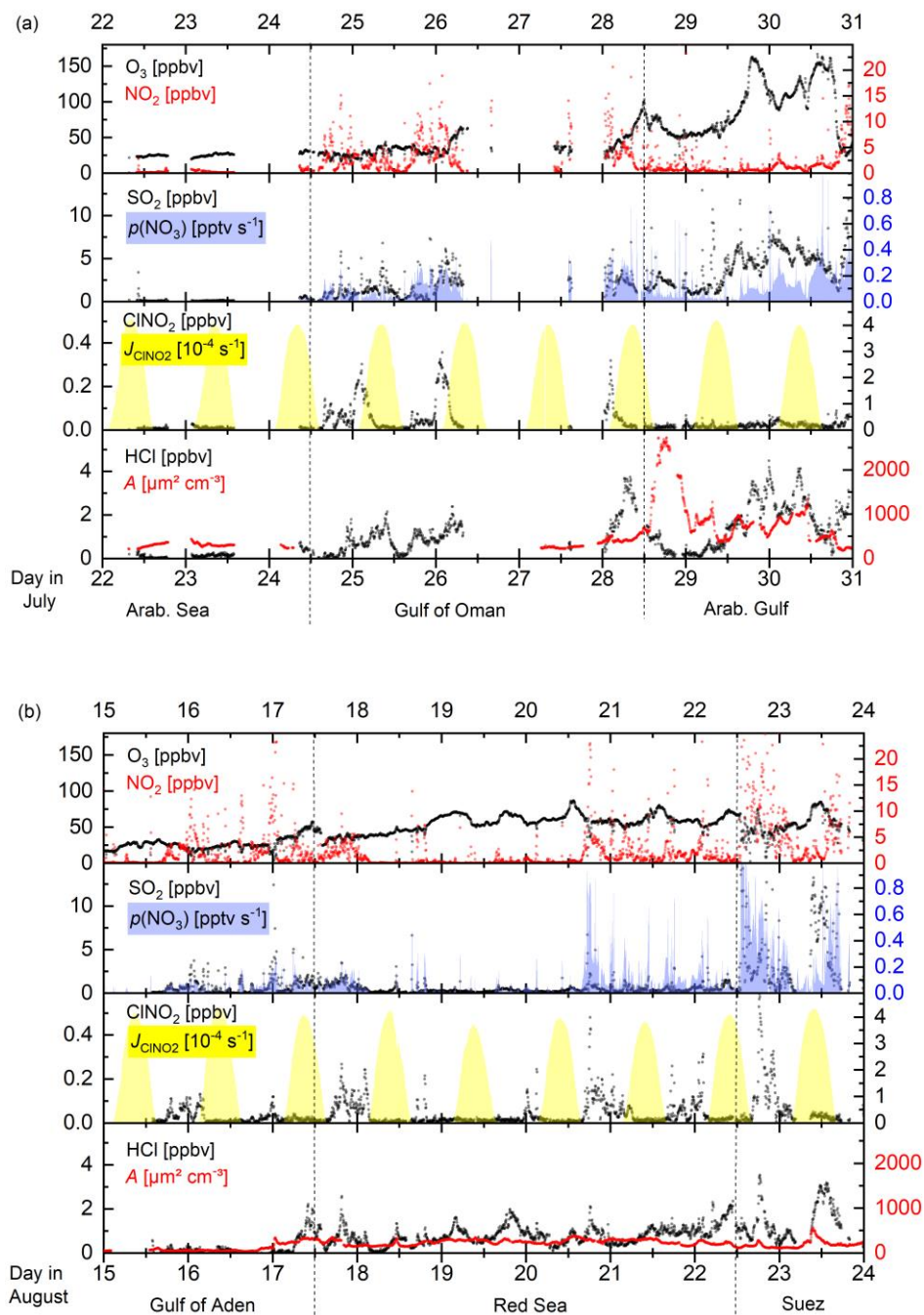
Figure 3: Time series of ClNO2 and trace gases related to its production (NO2 and O3) as well as the NO3 production rate, $p(NO_3)$, ClNO2 photolysis rate (J_{ClNO2}), PM1 particle surface area concentration (A) and HCl mixing ratio in different regions (separated by the dashed lines) for 9-day periods in the (a) first and (b) second leg.
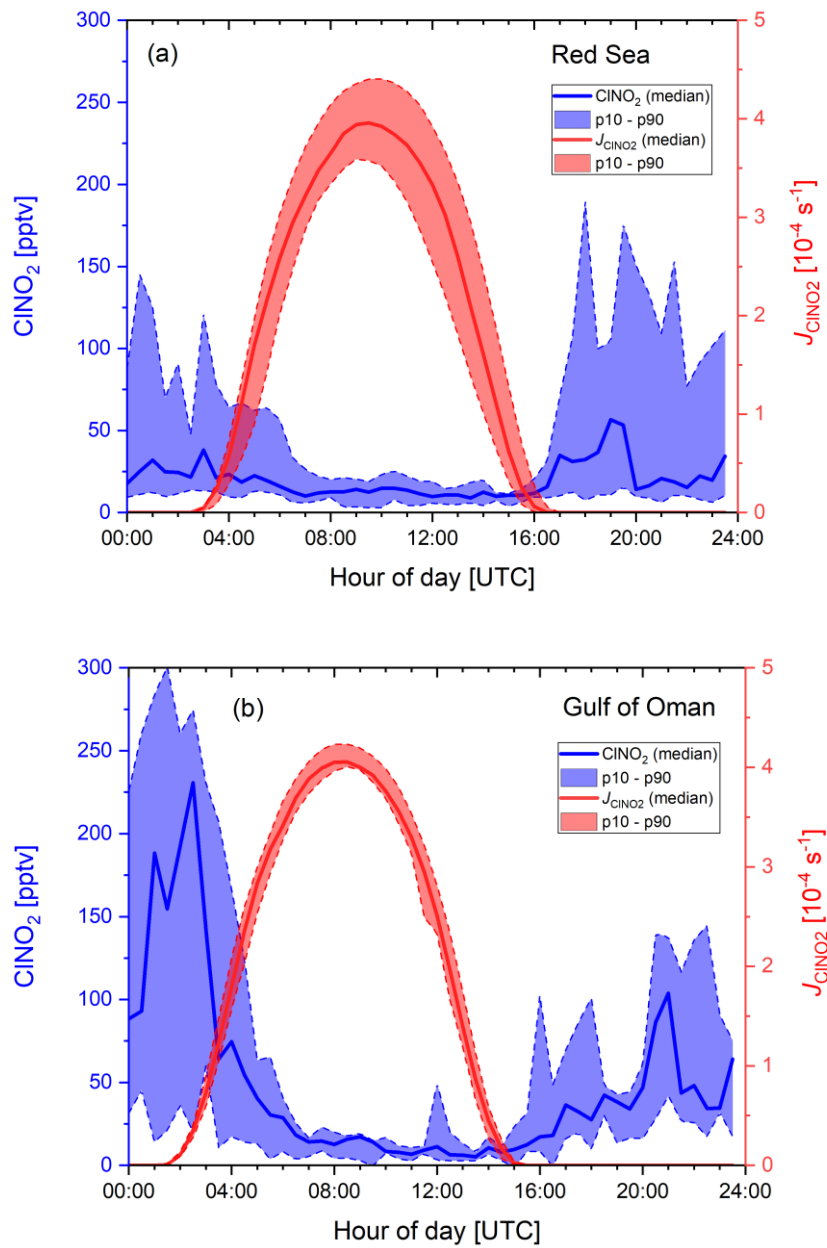
Figure 4: Diurnal profiles of ClNO₂ and its photolysis rate constant, J_{ClNO2}, for (a) the Red Sea and (b) the Gulf of Oman. The solid lines represent the median values, the shaded areas correspond to the 10th and 90th percentiles.
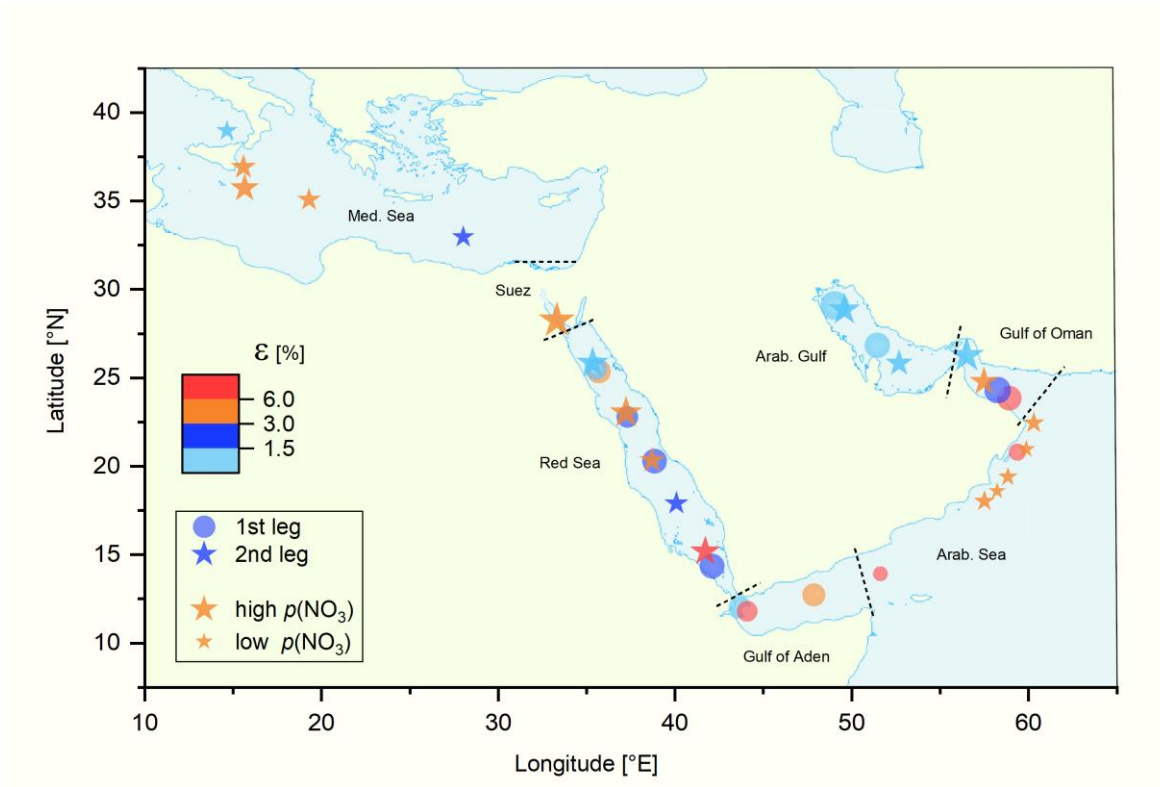
Figure 5: Median night-time values of ε (the ClNO$_2$ yield per NO$_3$ molecule formed) calculated via Eq. (4). The size of the symbols represents (logarithmic scale) the median NO$_3$ production rate, p(NO$_3$), ranging from 0.001 pptv s^{-1} in the Arabian Sea to 1.2 pptv s^{-1} in the Suez Canal. Different regions are separated by the dashed black lines.
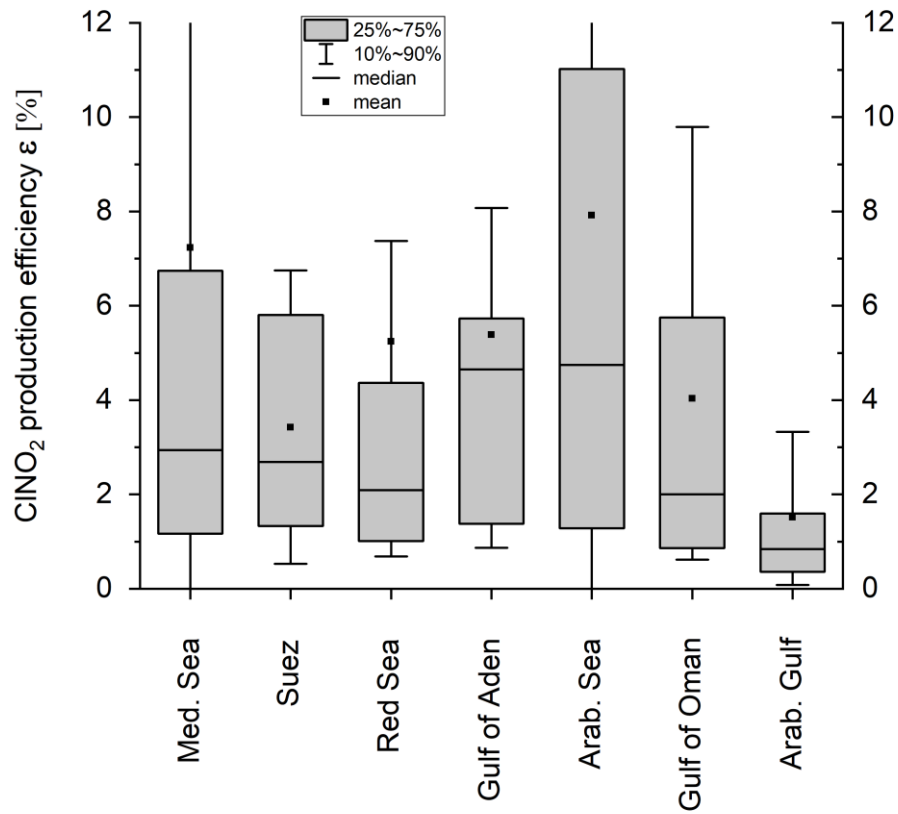
Figure 6: Box plots of ε (ClNO₂ production efficiency) for different regions, based on all corresponding individual night-time values calculated from Eq. (4).
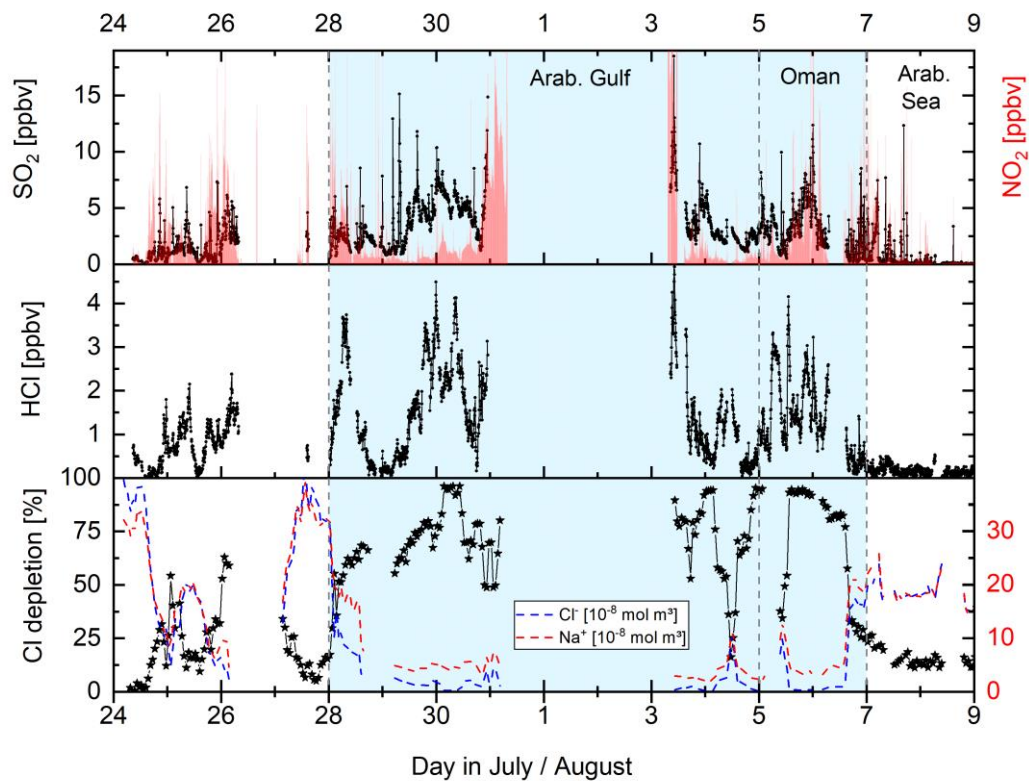
Figure 7: Co-variance between mixing ratios of SO_2, NO_2 and HCl and particulate chloride depletion (calculated from Eq. 12) illustrated by the difference in Cl^- and Na^+ (PM$_1$) measured. Chloride depletion of up to 90 % indicates effective acid displacement of HCl by HNO_3 and H_2SO_4 in this region.
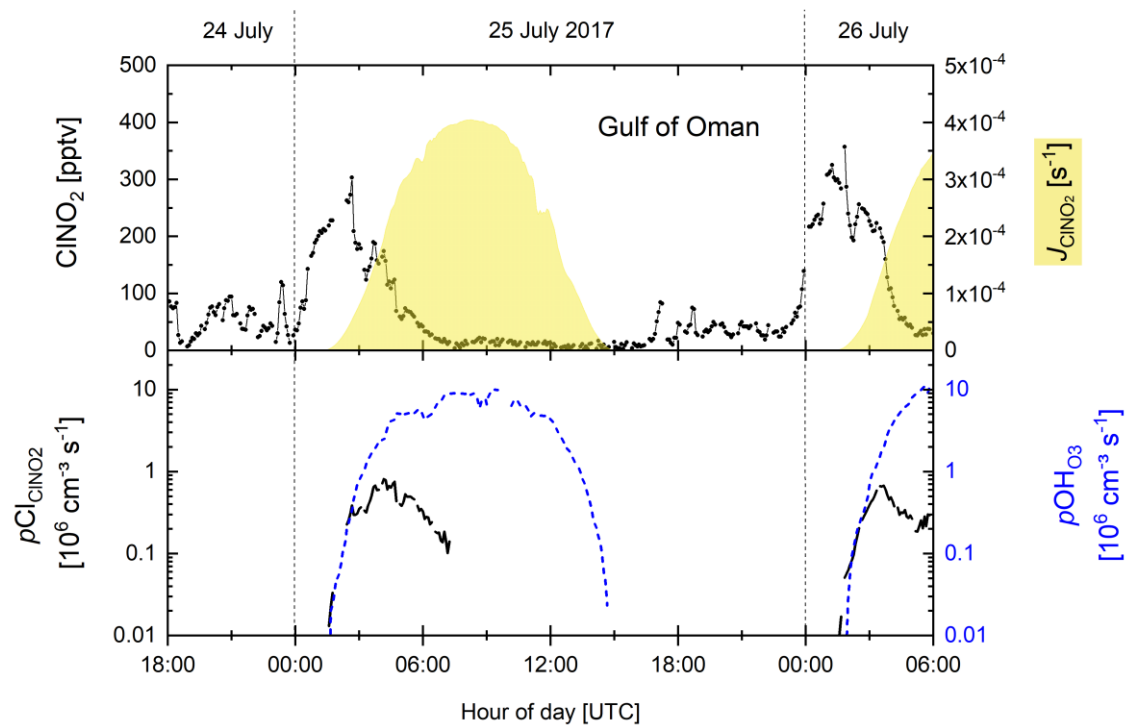
Figure 8: Time series of $ClNO_2$ mixing ratios, J_{ClNO2} photolysis rates and production of Cl-radicals from $ClNO_2$ photolysis (pCl_{ClNO2}) and OH-radicals from O_3 photolysis in the presence of H_2O (pOH_{O3}) for two consecutive nights in the Gulf of Oman.
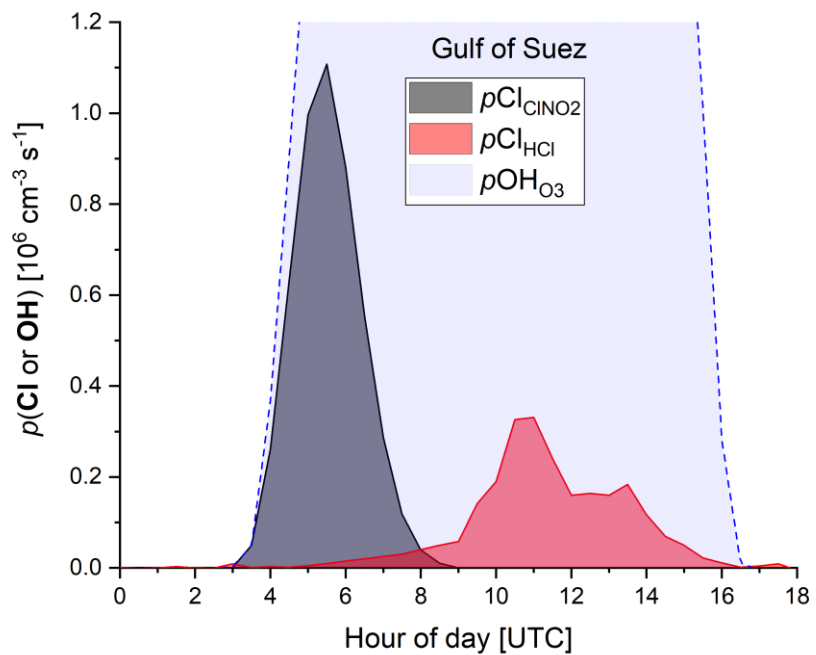
Figure 9: Production of Cl-atoms (from $ClNO_2$ photolysis and HCl + OH) and OH-radicals (from O_3 photolysis) over one diurnal cycle in the Suez Canal / Gulf of Suez.

Supplement of

Shipborne measurements of ClNO$_2$ in the Mediterranean Sea and around the Arabian Peninsula during summer

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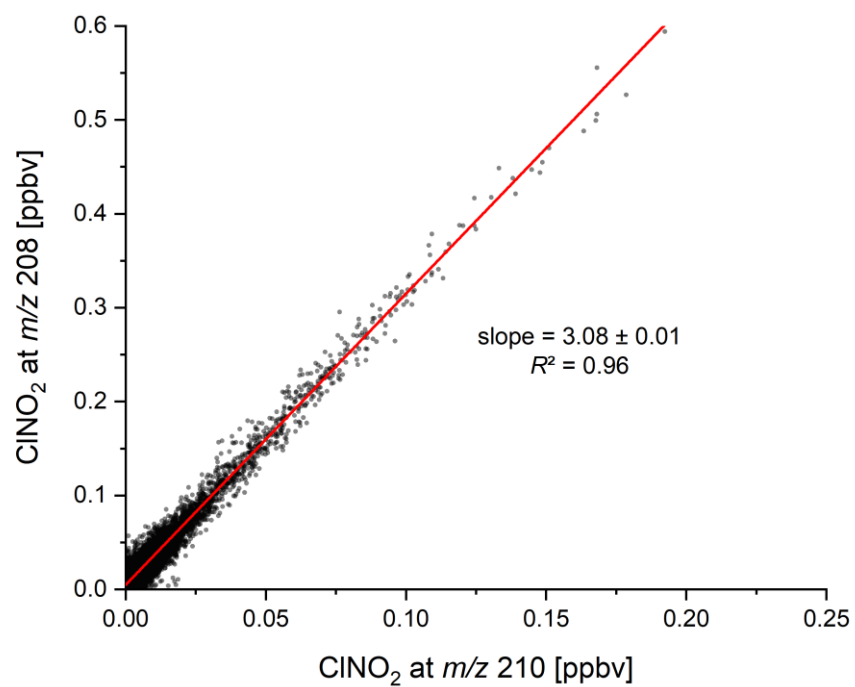
Figure S1: Correlation between $ClNO_2$ signals measured at *m/z* 208 and *m/z* 210.

Corrections to the aerosol particle surface area concentration

To derive the ambient PM_1 particle surface area concentration A from the measured dry particle surface area concentration (AMS) we applied a hygroscopic growth factor based on ambient RH and PM_1 aerosol composition. From the AMS measurements we derived the molar ratio of sulphate to ammonium of ≈ 2.2 (see Fig. S1), which is close to the ratio of ≈ 2 for completely neutralised ammonium-sulphate aerosol, indicating its dominance in the fine mode during AQABA. For most of the campaign $(NH_4^+)_2(SO_4^{2-})$ contributed 80-100 % to the total inorganic PM_1 aerosol mass (see Fig. S2) while nitrate and chloride were less abundant. A relative humidity-dependent growth factor G_{amsu} for the particle diameter was calculated using the parameterisation (Eq. S1) of Lewis (2008) for pure ammonium-sulphate drops with a = 0.78 and b = 1.90.

$$G_{amsu}(RH) = a \left(b + \frac{1}{1-RH} \right)^{1/3} \tag{S1}$$

The growth factor was calculated considering the RH of the sample air after passing the aerosol dryer (see Sect. 2.5 in the manuscript) which was usually in the range 40 ± 10 %. To account for the organic mass fraction we calculated the volume fraction v_{org} of organics using a typical density of 1400 kg/m³ for oxidised organics in aged atmospheric aerosol (Gysel et al., 2007) and the volume fraction v_{amsu} of $(NH_4^+)_2(SO_4^{2-})$, assuming that all NH_4^+ and SO_4^{2-} measured originated from ammonium-sulphate. In the next step we derived a combined growth factor by applying a mixing rule (Eq. S2) where the water activity a_w was approximated by the relative humidity (Gysel et al., 2007):

$$G_{mixed} \approx \left(\sum_i v_i \, G_i^3 \right)^{\frac{1}{3}} \tag{S2}$$

For the organic fraction a growth factor of $G_{org} = 1.20 \pm 0.10$ at $a_w = 0.9$ and a ratio of $\frac{G_{org}}{G_{amsu}} \approx \frac{1.20}{1.80}$ was reported by Gysel et al. (2007). Assuming a similar RH-dependence for G_{org} than for G_{amsu} results in Eq. S3 which can be inserted into Eq. S2 along with Eq. S1.

$$G_{org} \approx 1 + \frac{1}{4}(G_{amsu} - 1) \tag{S3}$$

The resulting surface area growth factor $G_A = (G_{mixed})^2$ as a function of RH is shown in Fig. S3. The particle surface area concentration (PM_1) reported in Table 1 of the manuscript already includes this correction. The distribution of surface area growth factors for the whole AQABA campaign is shown in Fig. S4 and has a campaign average of 1.32 ± 0.24.
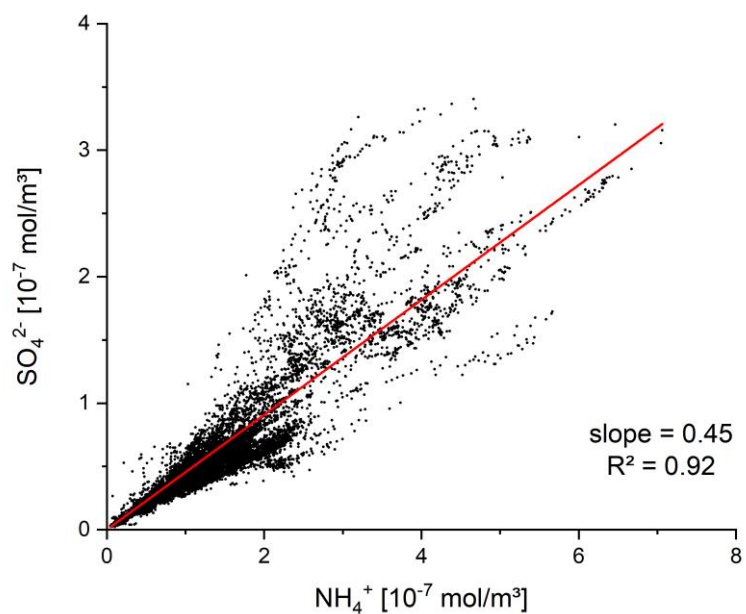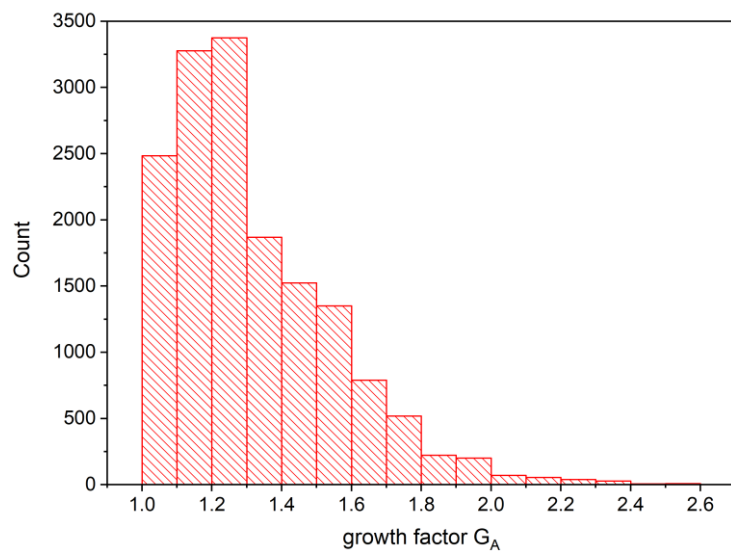
Figure S2: Correlation between PM_1 (AMS) ammonium and sulphate (in mol m^{-3}).
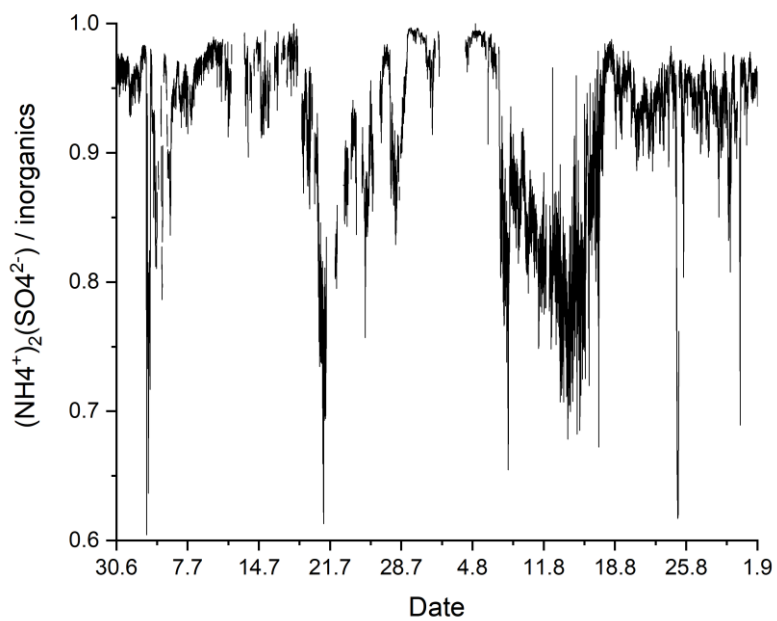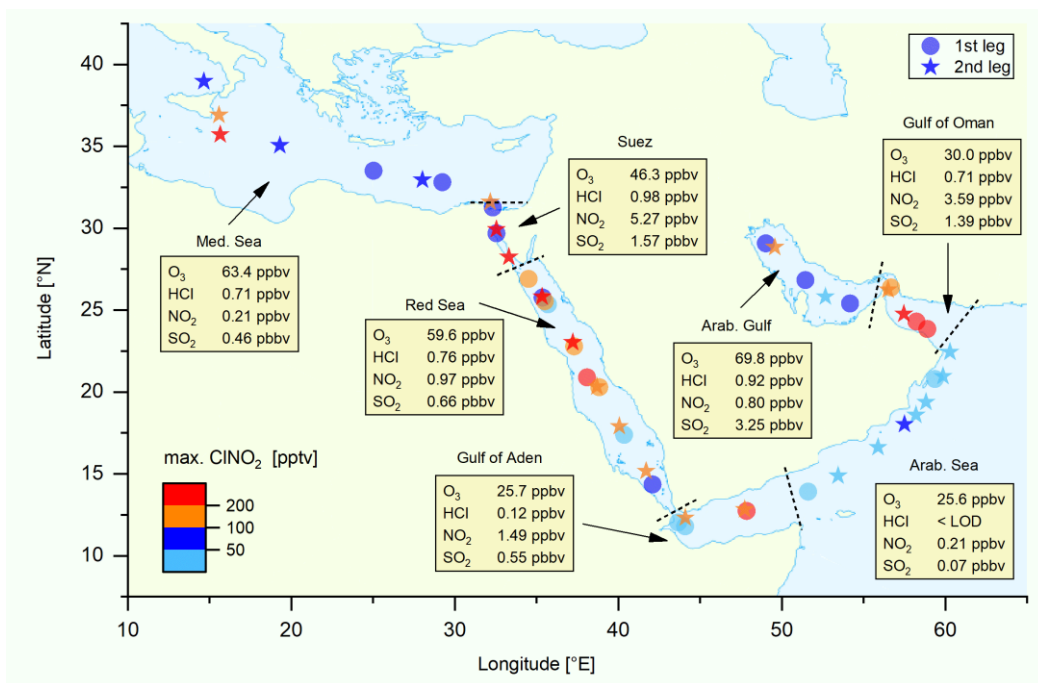

Figure S3: Contribution of ammonium-sulphate to the total non-refractory inorganic PM_1 aerosol mass.
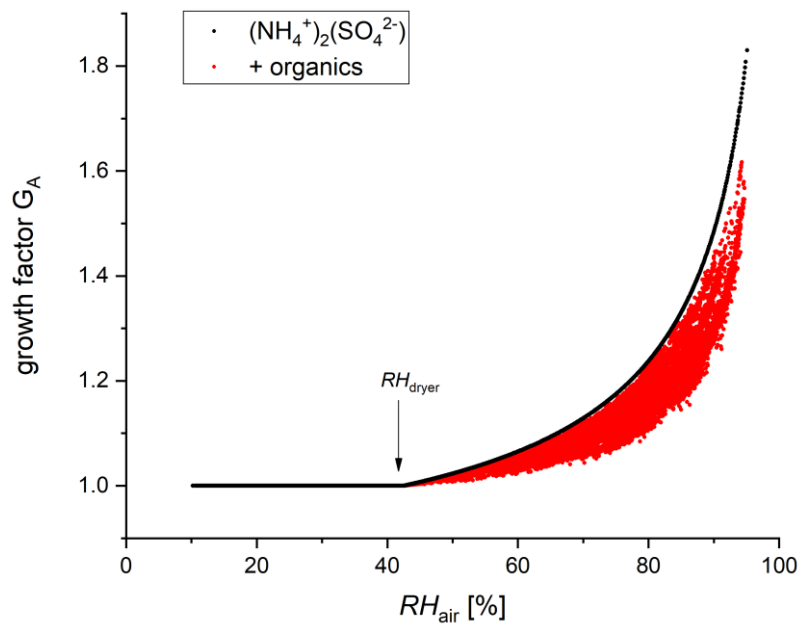
Figure S4: Surface growth factors for pure ammonium-sulphate (black) and for a mixture with organics (red).


Figure S5: Distribution of calculated surface area growth factors G_A for the whole AQABA campaign.

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Figure S6: Map of maximum $ClNO_2$ mixing ratios on individual nights together with (median) night-time mixing ratios of O_3, HCl, NO_2 and SO_2 for different regions demarked by dashed lines. Circles and stars represent data obtained on the first and second legs, respectively.
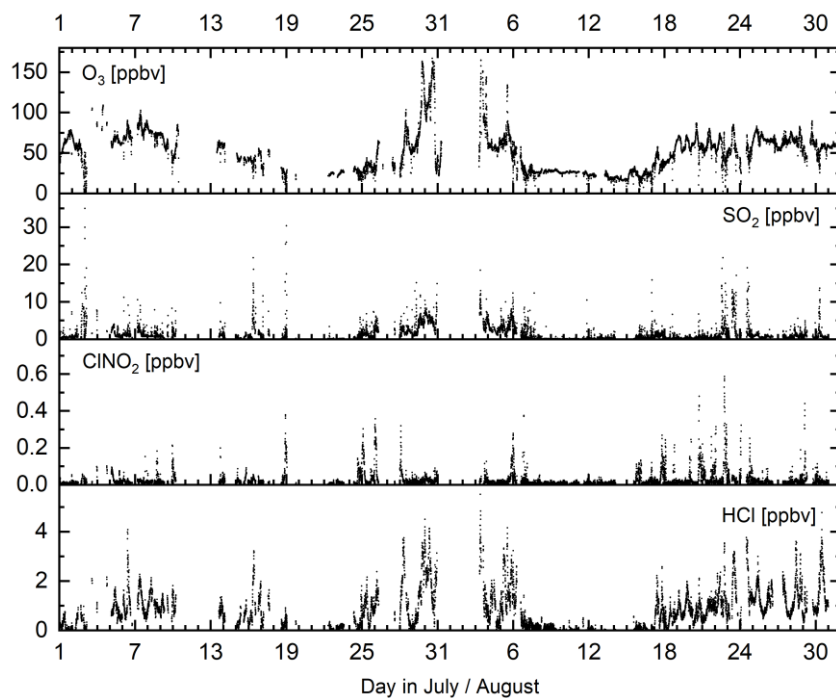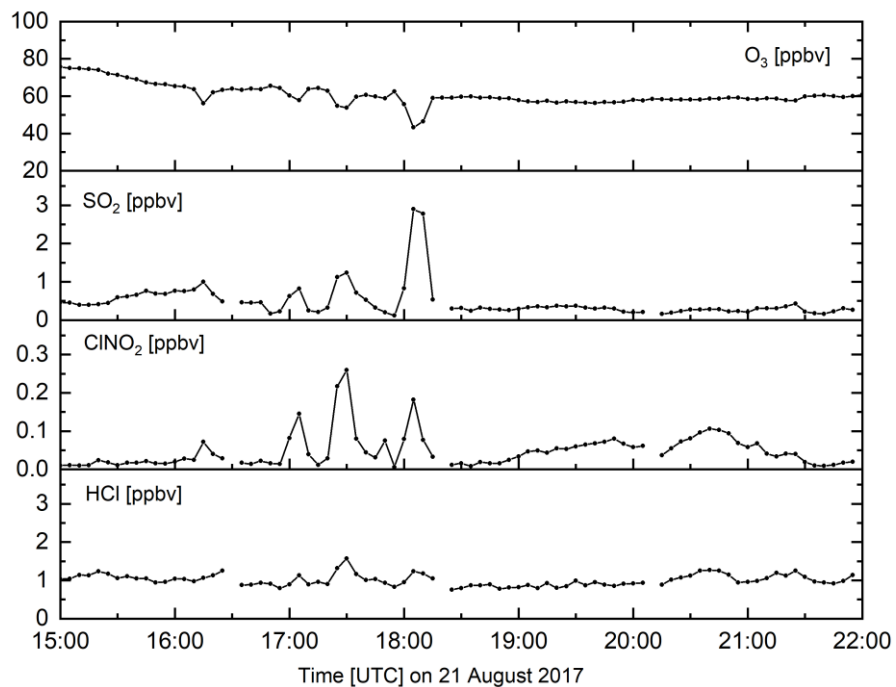

5   **Figure S7:** Time series of O_3, SO_2, $ClNO_2$ and HCl mixing ratios during the AQABA campaign.
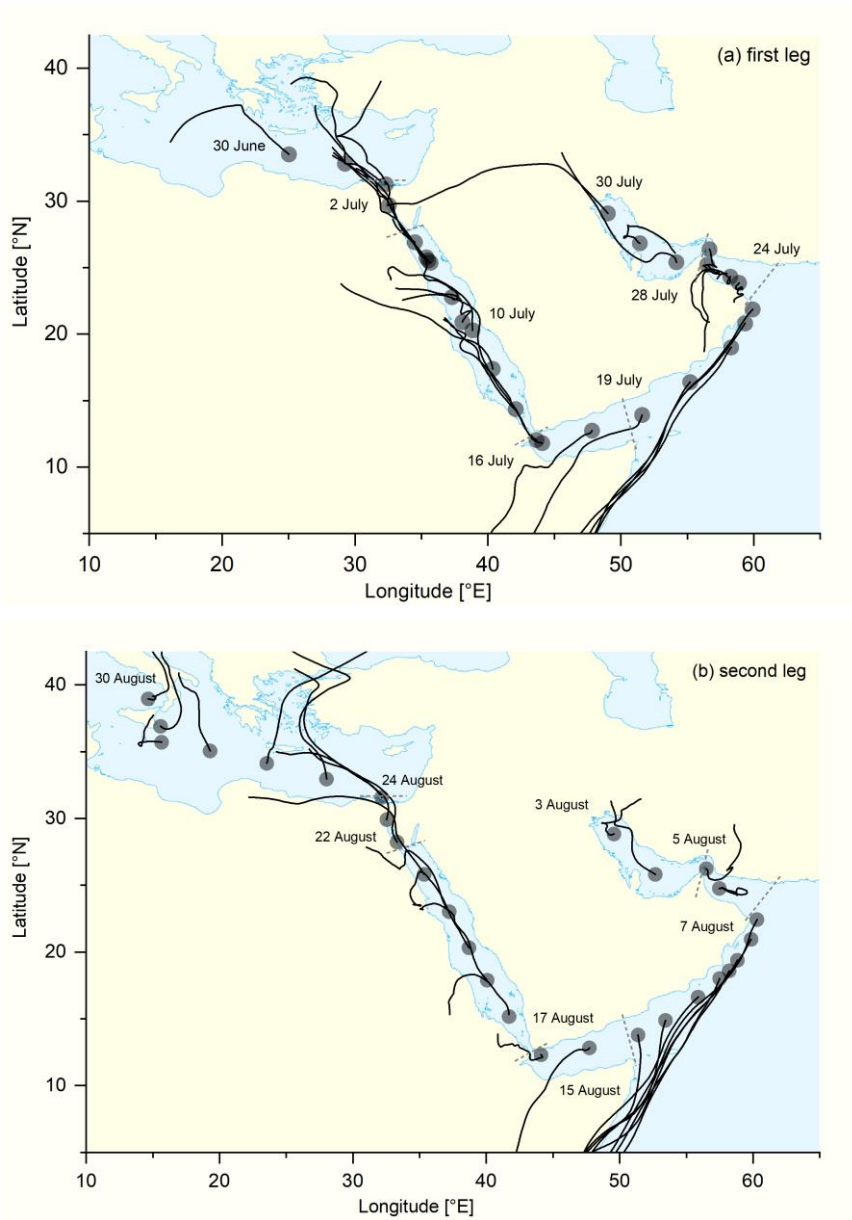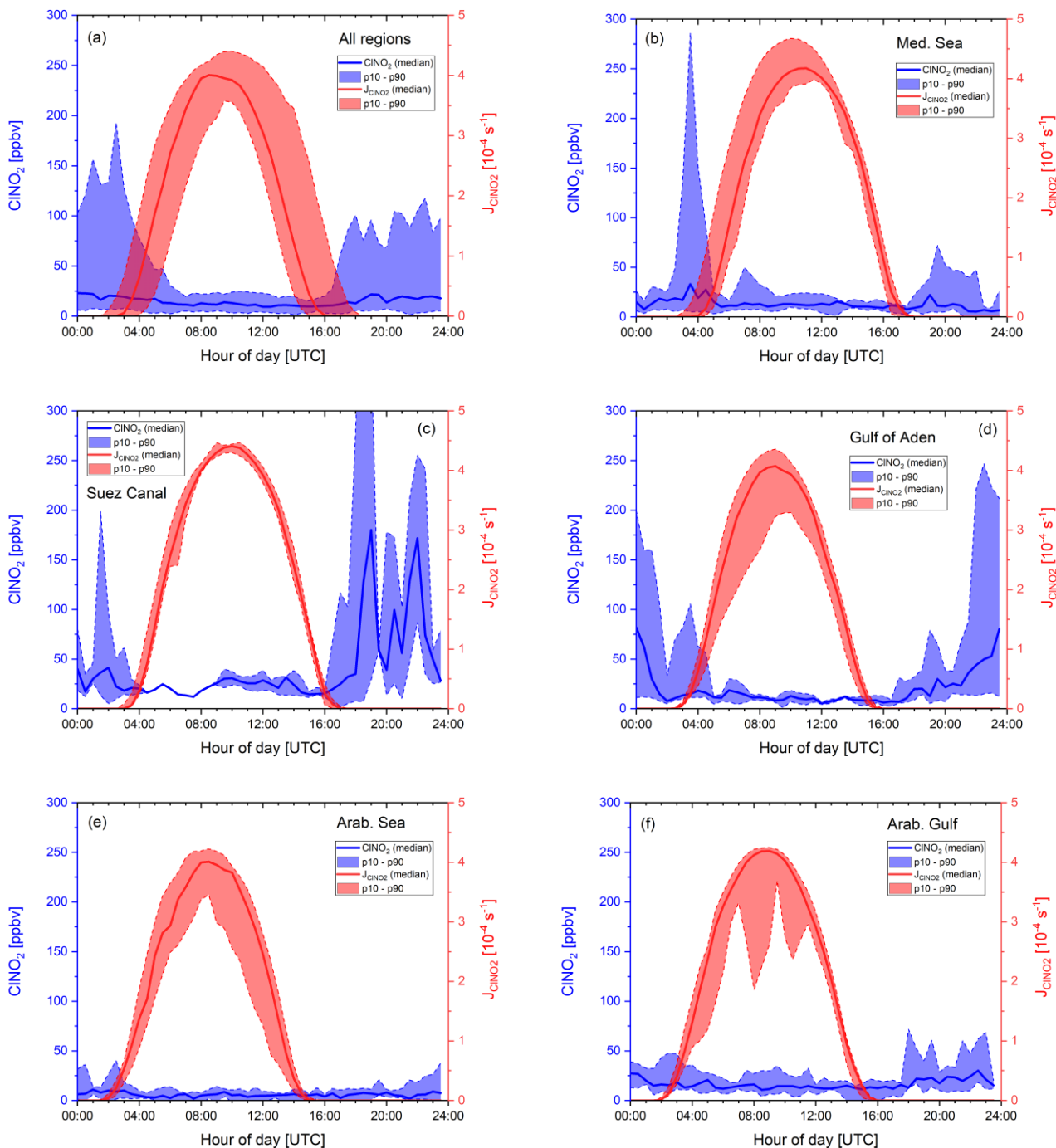
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Figure S8: 48-h back-trajectories calculated with HYSPLIT (at 100 m above sea-level) representative for the corresponding nights.

Figure S9: Observation of ClNO$_2$ in distinct ship plumes between 16:30 and 18:30 UTC (indicated by increased SO$_2$ mixing ratios and O$_3$ titration due to NO) on the 21st of August in the Red Sea.

Figure S10: Diurnal profiles of ClNO₂ for (a) all regions combined, (b) the Mediterranean Sea, (c) the Suez Canal and Gulf of Suez, (d) the Gulf of Aden, (e) the Arabian Sea, and (f) the Arabian Gulf. p10 and p90 correspond to the 10th and 90th percentiles.

Additional details to the derivation of the ClNO₂ production efficiency ε

To investigate the uncertainty related to assumptions made in the calculation of ε, we define five different methods A-E and

5 compare them to each other by applying them to the whole AQABA dataset (Fig. S9).

Method A: The starting point t_0 for NO_3 formation is set to sunset minus 40 ± 10 min as described in the manuscript but all data points before sunset were excluded from the analysis due to a large uncertainty in reaction time. This way 4175 values of ε were obtained throughout the campaign. In *Method A*, we inherently assume that the air mass we probe has not been influenced by fresh NO emissions since the beginning of the night. A fresh NO emission would mean that the air mass at the

10 beginning of the night would have contained less NO_2 than calculated in Eq. 3 resulting in an overestimation of the integrated NO_3 production and thus values of ε calculated by *Method A* represent a lower limit.

Method B: To account for fresh emissions of NO (e.g. by passing ships), the reaction time t' was calculated from Eq. 5 where s represents the number of NO_2 molecules required to make NO_y and is 1 when NO_3 reacts directly with VOCs and 2 when NO_3 reacts with NO_2 to form N_2O_5, which subsequently hydrolyses to HNO_3. As discussed later, the direct NO_3 losses are

15 dominant throughout the campaign compared to the heterogeneous N_2O_5 production, so to a good approximation, $s = 1$. As discussed by McDuffie et al. (2018) inherent to the use of this expression is the assumption that NO_y is conserved during the night; any losses of NO_y (e.g. via deposition of HNO_3) leading to an underestimation of the true reaction time. Whenever t' is shorter than the time elapsed since sunset (see *Method A*), t' is used to integrate the NO_3 production term. As the calculated, night-time air mass age depends on the ratio between [NO_2] and [NO_y], the calculation breaks down whenever a fresh NO

20 emission (e.g. from a nearby ship) is injected into an air-mass and unreacted NO is still present. In this case the NO_2 to NO_y ratio would be decreased and the age of the air-mass overestimated. To avoid this, we only use data where NO is below the detection limit, resulting in a total number of $N_B = 2987$ data points.

Method C: Here we only consider data points where the calculated age of the air mass (as derived in *Method B*) is equal to or exceeds the time elapsed since sunset as derived in *Method A*. These air masses are unlikely to have been impacted by recent

25 emission. As loss of NO_z by deposition will result in an air mass age that is shorter than the true one, we relax the criterion for equality of reaction times by also including calculated air mass ages that are up to 25 % shorter (i.e. $t' \geq 0.75\,(t - t_0)$). The resulting dataset is reduced to $N_C = 1742$ data points. All values of ε discussed in the manuscript were derived using method C.

Method D: ClNO₂ mixing ratios close to the limit of detection (LOD) result in a higher uncertainty and variability in ε,

30 especially when the NO_3 production term is also small. In *Method D*, we use only data in which the ClNO₂ mixing ratio is at least 25 pptv, which represents the LOD + the median, daytime HCl-interference (Sect. 3.1). This drastically reduces the size of the dataset to $N_D = 280$ data points.

Method E: In method E, we examine the efficiency of ClNO$_2$ formation only during a few nights when its mixing ratio exceeded 100 pptv. The intention here is not to derive e.g. a regional mean value, but to indicate that even when biasing the dataset to apparently efficient ClNO$_2$ generation, ε remains low. In Method E, only 50 data points remain (1.2% of the dataset analysed in *Method A*).

5   We summarise values of ε as median and mean values for the entire campaign in Fig. S9, segregated into the five different methods used to select data and derive the reaction time. As described earlier, *Method A* can be understood as a lower limit for ε providing a median efficiency of only 1.4 % with a range from 0–8 % (10th and 90th percentiles) and a large difference between mean and median values. More reliable median values of ε = 2.8 % and 2.7 % are provided by *Methods B and C*. *Method D*, results in identical median and mean values of ε = 4.5 %, although a bias towards higher values is difficult to rule

10   out as low ClNO$_2$ mixing ratios were excluded. For *Method E*, where we only consider data with ClNO$_2$ > 100 pptv, we derive a larger median value of ε = 6.4 %. If we consider only the individual maxima in the ClNO$_2$ mixing ratio (above 100 pptv) on any particular night, we derive 17 values of ε that vary between 1.1 and 11.2 % which are listed in Table S1.

In our manuscript (Fig. 5 and 6) we exclusively report values of ε derived by method C. In comparison to Fig. 6 in the manuscript, Figure S10 illustrates how the box-plots for the seven regions would be altered when applying method D instead

15   of method C. The variability is generally decreased and median values are shifted towards higher values, but the general picture and conclusions are not changed.
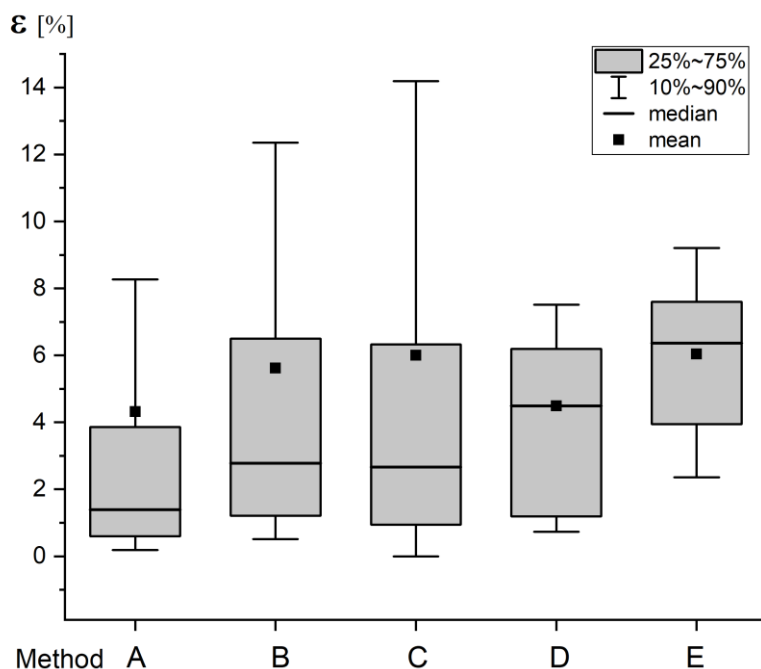
**Figure S11:** Box plot of ε derived by the different methods A–E for the whole AQABA campaign. The whiskers represent the 10th and
20   90th percentiles.
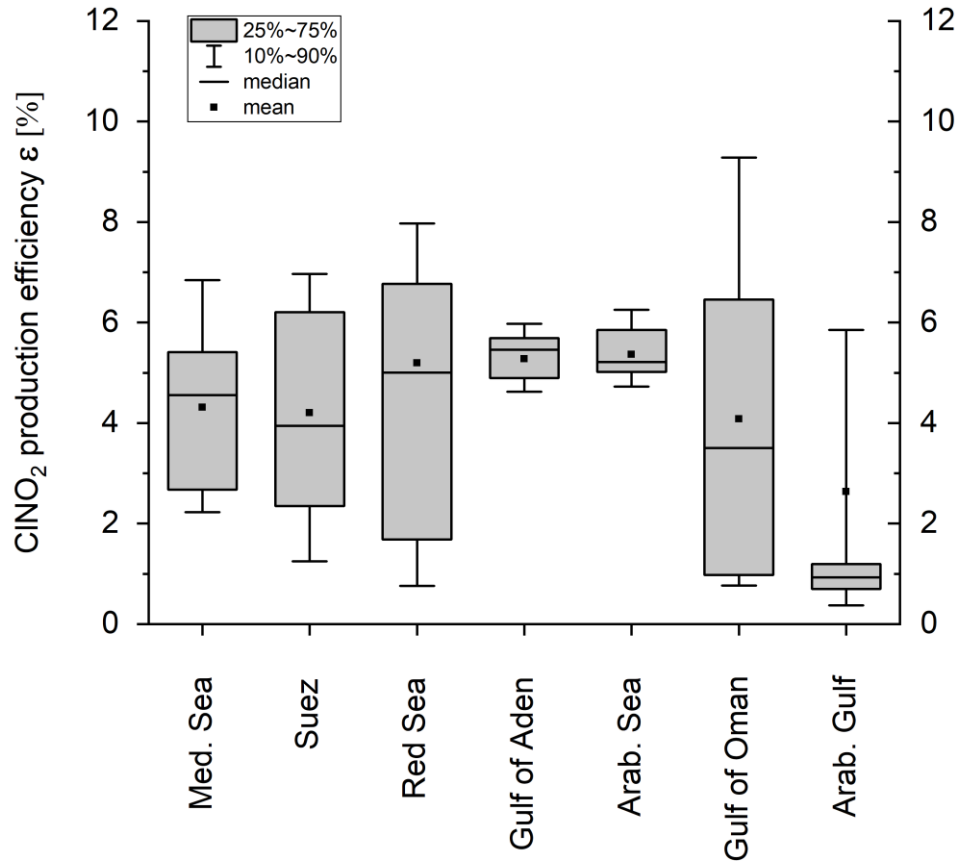
Figure S12: Median values of ε (ClNO₂ production efficiency) for each region, calculated from individual night-time values (between 10 and 104 per region) based on Eq. 4 in the manuscript but using method D instead of method C.

5
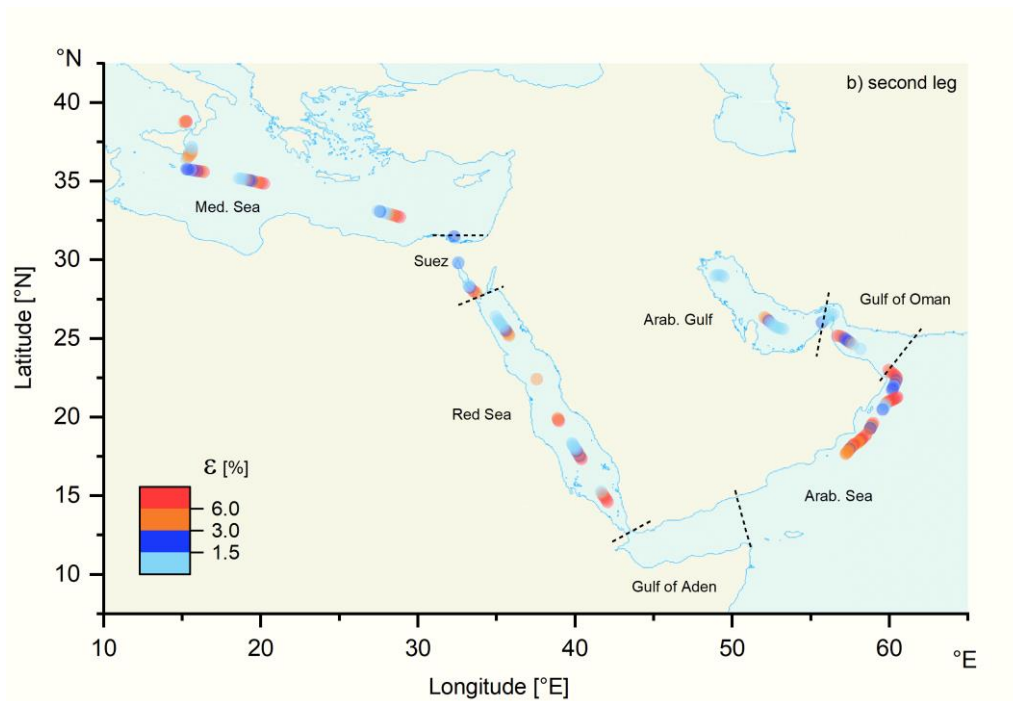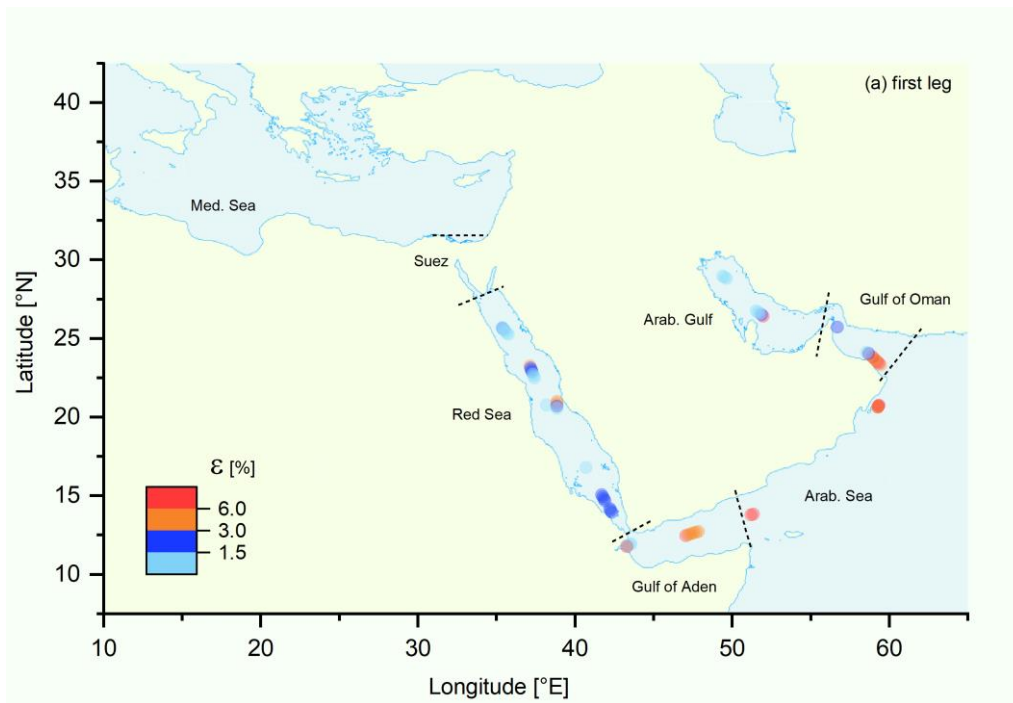
Figure S13: 1-hour averaged values of ε along the ship track for (a) first and (b) second leg, calculated via Eq. (4).

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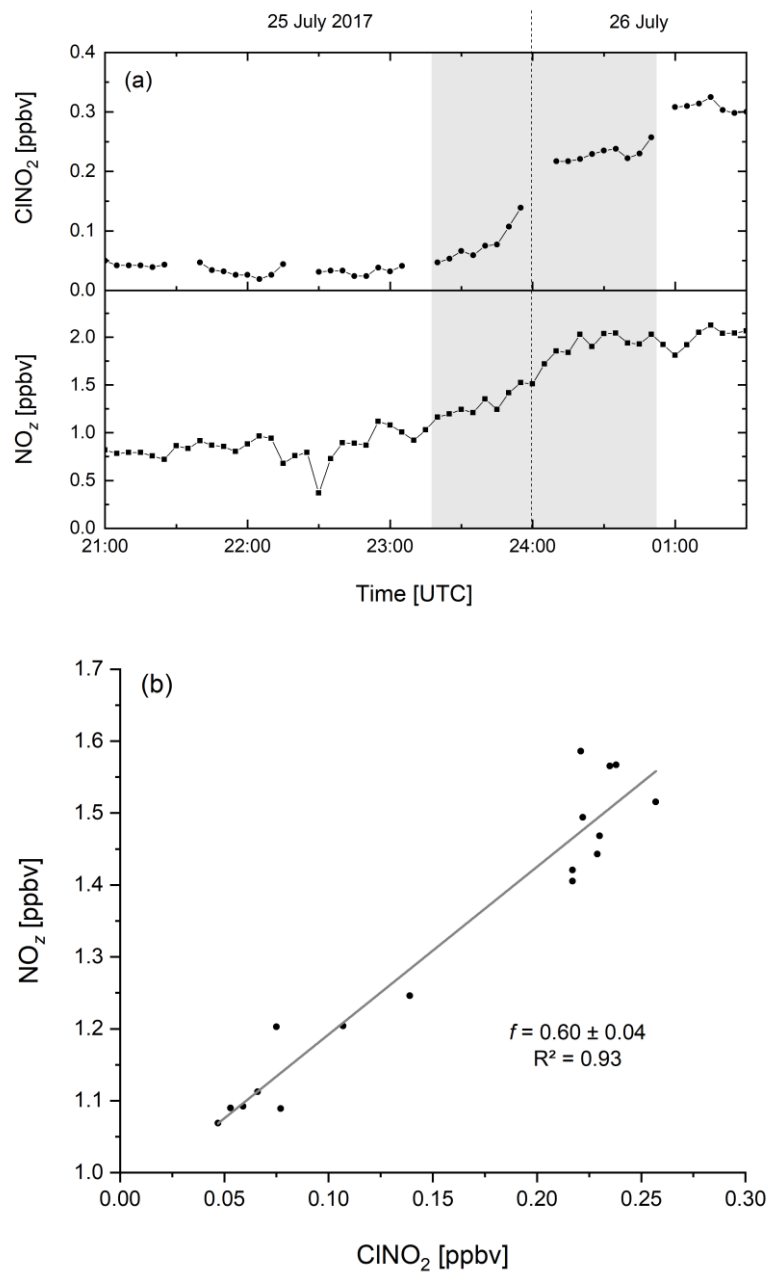
Figure S14: (a) Time series of ClNO₂ and NO_z on 25–26 July 2017 in the Gulf of Oman. (b) The slope of NO₂ vs. ClNO₂ can be used (Eq. 9) to calculate $f = 0.60 \pm 0.04$ for the ~ 2 h period (grey shaded area in Fig.S14a).

Table S1: Observed ClNO$_2$ plumes [1] (typical duration of one to several hours) with mixing ratios above 100 pptv.

Date, time [UTC]	Region	t [h]	t' [h]	ClNO$_2$ [pptv]	N$_2$O$_5$ [pptv]	NO$_{3, int}$ [ppbv]	ε [%]	O$_3$ [ppbv]	SO$_2$ [ppbv]	HCl [ppbv]	NO$_2$ [ppbv]
08.07. 17:50	Red Sea	1.8	1.5	128	-	3.17	4.0	61.5	3.0	1.1	6.9
24.07. 16:05	Oman	1.8	2.3	143	-	1.28	1.1	25.7	1.1	0.3	7.0
03.08. 18:00	Arab. Gulf	2.9	2.6	115	53	9.65	1.2	78.8	3.0	1.8	6.7
06.08. 18:45	Oman	4.3	3.9	173	-	1.55	1.1	28.1	1.4	0.7	2.9
06.08. 20:35	Oman	6.1	6.9	159	-	2.86	5.6	20.0	4.0	0.6	5.2
17.08. 18:00	Red Sea	3.0	4.5	120	9	1.54	7.8	34.4	1.6	0.7	3.2
17.08. 19:25	Red Sea	4.4	3.6	212	14	3.03	7.0	34.7	2.1	1.8	3.9
17.08. 20:05	Red Sea	5.1	4.7	168	8	2.03	8.3	32.0	1.3	1.5	2.4
18.08. 17:00	Red Sea	1.8	1.5	106	23	1.49	7.1	42.9	0.9	0.6	4.6
20.08. 16:45	Red Sea	1.3	1.0	144	59	1.90	7.6	58.7	0.7	1.0	5.8
21.08. 17:05	Red Sea	1.5	1.4	145	53	2.13	6.8	57.9	0.8	1.1	5.9
21.08. 17:30	Red Sea	1.9	2.5	259	87	3.14	8.2	53.8	1.2	1.6	7.1
21.08. 23:40	Red Sea	8.1	8.5	103	27	9.78	1.1	58.0	0.7	1.1	2.3
22.08. 16:50	Suez	1.0	1.3	123	118	1.96	6.3	45.7	2.5	1.2	11.4
22.08. 18:20	Suez	2.5	2.1	586	355	7.09	8.3	56.4	4.6	3.5	11.3
22.08. 21:40	Suez	5.8	4.6	308	52	8.77	3.5	39.4	1.2	1.2	7.6
22.08. 22:30	Suez	6.7	6.1	222	47	9.45	2.3	43.0	1.0	1.1	5.9

[1] Regions: Red Sea, Gulf of Oman (Oman), Arabian Gulf (Arab. Gulf) and Suez Canal / Gulf of Suez (Suez). t denotes the time since sunset; t' corresponds to the air mass age calculated from Eq. 5 in the manuscript NO$_{3,int}$ is the total amount of NO$_3$ produced over the course of the night and ε is the ClNO$_2$ production efficiency (Eq. 4 in manuscript).

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