The following contains the comments of the referee (black), our replies (blue) indicating changes that will be made to the revised document (red).

Reviewer #1

This is an excellent and very careful study of the kinetics of the reaction of OH radical with NO2, a reaction of central atmospheric significance. The main finding associated with this work (the second paper from this group on this reaction) is the finding of a significant enhancement to the OH/NO2 rate coefficient in the presence of water vapor. This enhancement is quantified in both N2 and He bath gas, and as a function of temperature. The data appear to be of the highest quality, previous data are discussed in detail, and the modeling work adds significantly to the atmospheric context, and hence the overall value of the manuscript. Overall a superb study! I recommend publication in ACP, after consideration of the minor comments listed below.
We thank the reviewer for the careful review and the positive assessment of our manuscript.

L12 – molecule misspelled. Corrected

L64 – exponent should be +3. Corrected

L127 – evaluation of atmospheric. Corrected

L148- Maybe it is obvious, but it could be added to the text that fitting with an increased Fc gives a lower ko, (which doesn't fit the data).
We have added:
We note that using a higher F_c=0.39 resulted in a lower value of k_0^{He} equal to 1.0×10^{-30} cm^6 molecule^{-2} s^{-1}.

Page 4,5 - The pure He data are taken here at 292 K, somewhat below other 'room temperature' datasets that comparisons are being made with. Has this has been taken into account? If not, could this account for at least some of the systematic discrepancies with D'Ottoni or Morley?
In Figure 1, we did not attempt to apply any corrections to previous works to account for temperature differences. In the more rigorous comparison of our work with others (Fig S3-S5), we used our parametrization at the given temperatures of the previous studies.
We now write in the Figure S3, S4 and S5 captions the temperature for which *k(T,p)* was calculated.

On a related point, there is a T-dependence to the He ko that appears in various places (captions to Fig 2, Fig 4b, L252). What is the origin of this value? (I think it is mostly unnecessary).
The T-dependence in He was determined in our laboratory over a small range of temperatures (277, 292 and 332 K) and pressures. We have added the data points used to determine *m*(He) in Table 1, a Figure (S10) in the supplementary information and the text below:

The T-dependence factor in He, *m*(He), was determined to be 3.1 over the temperature range from 277 to 332 K (Table 1 and Figure S6).

T (K)	*p* (Torr)	M[a]	OH precursor	k_1[b]
277	48.6	1.68	H_2O_2	1.59 ± 0.14
292	25.1	0.83	H_2O_2 [c]	0.75 ± 0.07
	50.0	1.65	H_2O_2	1.37 ± 0.08
	75.1	2.47	H_2O_2	1.88 ± 0.12
	102.9	3.39	HNO_3	2.32 ± 0.15
	206.9	6.81	HNO_3 [d]	3.73 ± 0.25
	300.7	9.89	HNO_3	4.64 ± 0.29
	405.8	13.35	HNO_3	5.54 ± 0.37
	495.6	16.30	HNO_3	6.29 ± 0.40
	595.0	19.57	HNO_3	6.83 ± 0.42
	689.1	22.67	HNO_3	7.46 ± 0.46
332	28.1	0.82	H_2O_2	0.60 ± 0.06
	56.8	1.65	H_2O_2	0.99 ± 0.08
	85.4	2.48	H_2O_2	1.34 ± 0.10

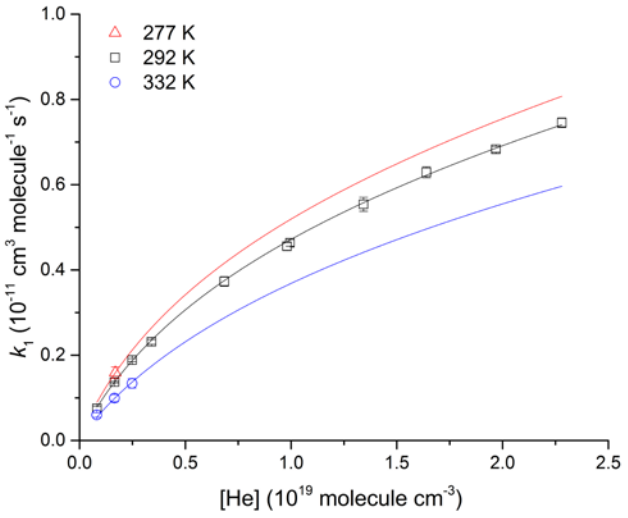

Figure S6. Values of k_1 as a function of He concentration at 277, 292 and 332 K. Errors are 2σ statistical only. The solid line is a fit to our data using Eqn. (4) with $k_0 = 1.4 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-1} s^{-1}, $F_c = 0.32$, $m = 3.1$ and $n = 0$.

L209- delete molecules Corrected
L252 – temperature dependence Corrected
L254 – delete use Corrected
L275 and forward – The tests for HONO and HNO3 are really nice !
L297 – temperature instead of temperatures Corrected
L342 – H2O, not HO2, as product of (R7) Corrected
L406 – short at Corrected
L408 – temperature and pressure conditions Corrected

The following contains the comments of the referee (black), our replies (blue) indicating changes that will be made to the revised document (red).

Reviewer #2

The manuscript describes the determination of the rate constant of the OH + NO2 reaction with He and N2 as bath gases in presence and absence of gaseous H2O. A quasi-static reaction cell was used, and OH was produced by pulsed laser photolysis of HNO3, H2O2, or O3/H2O mixtures. Pseudo-first order conditions with respect to [OH] were applied. The OH concentration was monitored time-resolved by laser-induced fluorescence, and the (crucial) NO2 concentration was carefully determined with two different absorption-spectroscopic approaches. A notable increase of the OH + NO2 rate constant in He and N2 when H2O is present was observed and associated with a particularly high efficiency of H2O for collisional stabilization of the HNO3 product. Non-linear mixing rules for the collisional efficiencies seem to apply. Very careful parameterizations and statistical evaluations of the experimental results, including earlier literature data, were performed and discussed in great detail, also with respect to branching between HONO2 and HOONO as reaction products. The newly parameterized rate constant is incorporated in a 3D chemical transport model, and effects on quantities such as the atmospheric HNO3/NO2 ratio, the atmospheric concentration of OH, or the HOONO/HO2NO2 ratio are assessed. All in all, this is a very nice paper bridging high-level state-of-the-art laboratory measurements with global atmospheric modeling. So the topic is at the very heart of ACP, and I recommend publication essentially ′as is′ with only very few, very minor points to be considered by the authors:

We thank the reviewer for the careful review and the positive assessment of our manuscript.

line 35:′gases′ should probably read ′gas′ Corrected

line 68:′O3-H2O′ should probably better read ′O3/H2O′ Corrected

Tables 1 and 2: please specify/explain M

In Table 1, we modified the caption below with:

Molecular density M(He) in units of 10^{18} molecule cm^{-3}

In Table 2, we modified in the caption,

Molecular density M(He-H_2O) or M(N_2-H_2O) in units of 10^{18} molecule cm^{-3}

Fig. 1, figure caption: please give the parameters m and n

The figure 1 caption now reads:

The solid line is a fit to our data using Eqn. (4) with $k_0 = 1.4 \times 10^{-30}$ cm^6 $molecule^{-2}$ s^{-1}, $k_\infty = 6.3 \times 10^{-11}$ cm^3 $molecule^{-1}$ s^{-1}, $F_c = 0.32$, m=3.1 and n=0.

Fig. 2, figure caption: please give the parameter n

The figure 2 caption now reads:

The black line is our parameterisation with $k_0 = 1.4 \times 10^{-30}$ cm^6 $molecule^{-2}$ s^{-1}, $k_\infty = 6.3 \times 10^{-11}$ cm^3 $molecule^{-1}$ s^{-1}, m = 3.1, n=0 and $F_c = 0.32$.

The following contains the comments of the referee (black), our replies (blue) indicating changes that will be made to the revised document (red).

Reviewer #3

This manuscript presents the first study of the effectiveness of H2O as a third body on the recombination of OH with NO2. This careful experimental study shows that water vapor is much more effective than N2 or O2 in causing recombination. It also presents a global modeling study of a new parameterization of the OH + NO2 reaction as compared to the IUPAC and JPL recommendations; this parameterization uses the results of a previous study showing that O2 and N2 have different efficiencies in quenching the products of OH + NO2 . The modeling suggests that HOONO could be a non-negligible reservoir of NOx in some parts of the atmosphere. This is a very important paper that is clearly in the scope of ACP. There are no major problems with the manuscript, but a few of points should be clarification or emphasized more strongly before publication in ACP.

We thank the reviewer for the careful review and the positive assessment of our manuscript.

My major concern about this manuscript is actually rather minor: In the global modeling, it is not clear how much of the affects of the new parameterization , occurs due to water vapor and how much due to the use of the results of the author's previous paper on N2 vs O2 as colliders. This should be made clear.

As we already mention, the impact of H_2O is limited to the boundary-layer (Fig. 5), above which the H_2O concentration decreases rapidly. We now emphasise this by writing:

The presence of water vapour does not impact on values of k_1 above the boundary layer.

The enhancement of the quenching of the energized HNO3 intermediate (HNO3 *) due to H2O vapor is presumably due to the strong hydrogen bonding between the two (stronger than OH-H2O or NO2-H2O). It would be good to make this explicit and add some references to the literature on the HONO2-HOH complex.

We have added text mentioning the HNO_3 – H_2O complex:

Water vapour is therefore a factor ≈ 8 more efficient than O_2, and a factor ~6 more efficient than N_2 as a quencher of the HO-NO_2 intermediate, which is qualitatively consistent with known strong binding (40 kJ mol^{-1}) in the HNO_3–H_2O complex (Tao et al., 1996).

There must be previous field work measuring [NO2]/[HONO2] and corresponding modeling work that did or did not find discrepancies. It seems that the authors should refer discuss a few of these, at least briefly.

This is an old problem in atmospheric chemistry. HNO_3 is not formed solely in the reaction between OH and NO_2 but also in heterogeneous processes that hydrolyse N_2O_5, the rate of which depends on poorly constrained factors such as e.g. the aerosol surface area. Uncertainties in modelled OH are large as are uncertainties for HNO_3 measurements, which in the boundary layer reflect deposition to surfaces, again poorly constrained. In short, uncertainties in kinetic parameters are only one factor that influence NO_2 and HNO_3 ratios and this issue is too complex to deal with properly in this manuscript.

Minor Issues:

Line 12: “molecule” is missing an “l” Corrected

Line 45: “being” should be “is” Corrected

Line 50-52. The sentence beginning “Theoretical calculations...” might better appear immediately after the discussion of the chaperone mechanism, rather than after the introduction of enhanced collider gases.
Corrected as suggested

line 68 “prevented” should be “preventing” Corrected

line 75: “in Tables 1 and 2.” might better be phrased as “in the notes to Tables 1 and 2.”
Corrected as suggested

It might help orient readers if the manuscript provided some idea of the conditions under which the OH + NO2 →HONO2 is nearly in the low-pressure limit and high-pressure limit.
We added the following text
Under the conditions of *T* and *p* relevant for atmospheric chemistry, the title reaction is in the fall-off regime.

Line 102: The manuscript states that the low vapor pressure of water prevents it from being used as a bath gas by itself, but 5 Torr of water vapor is roughly equivalent of 50 Torr of He. So I think that it would be clearer to say that it is not possible to determine k0 (H2O) by using pure water vapor as a bath gas.
This is correct. We prefer to remove this statement completely.

Lines 205 ff. “In other words....” It is not clear to this reader how it follows from the previous text that the total rate constant of a H2O-N2 bath gas is not the sum of individual rate constants k(Pi,T), where i=H2O or N2. It seems like a step of the logic has not been made explicit, and that it would help the reader if the manuscript made the logic clearer.
Starting the sentence with “In other words” was indeed confusing.
We have removed this.

In Section 3.2.1, it would be helpful to indicate the pressures at which k(P,300) deviates by more than 10% from the low-pressure and high-pressure limits. This would help orient the reader.
It is not obvious what insight this brings as there is no physical meaning associated with a 10% deviation. We believe this would be confusing and prefer not to follow this suggestion.

Line 298: the pressures only add up to 990 mbar, not 1 bar.
We now write:
The pressure of water vapour is 34 mbar, those of O_2 and N_2 are then 210 and 756 mbar, respectively

Lines 334 ff. If I understand correctly, the manuscript takes the branching ratio between reactions (1a) and (1b) from previous work. This is equivalent to assuming that water vapor

enhances both rate constants to the same extent. This assumption should be made very explicit in the manuscript.
We have added the following sentence to make this clear.
We assume α is independent of water vapour.

Lines 432-433: Are the H-OONO and H-OONO2 bond energies known from computational chemistry (well enough to determine which is stronger)?
We are not aware of such calculations and are not in a position to do them ourselves. The text simply introduces the possibility that the OH rate coefficients for H-OONO and H-OONO2 are different and suggests that theoretical calculations would be useful. Hopefully, someone will conduct such calculations and calculate the relative rate coefficient.

Figure 3 lacks error bars.
The error bars are already on the plot but are in most case smaller than the symbol. We have added to the caption:
With a few exceptions, the error bars (2 σ) are generally smaller than the symbols.

Table 2: The caption lists the range of [HOOH] used for the He-H2O experiments twice. I suspect one of these is for the N2-H2O experiments. Correction made.

Kinetics of the OH + NO_2 reaction: Effect of water vapour and new parameterisation for global modelling.

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Abstract. The effect of water vapour on the rate coefficient for the atmospherically important, termolecular reaction between OH and NO_2 was determined in He-H_2O (277, 291 and 332 K) and N_2-H_2O bath gases (292 K). Combining pulsed laser photolytic generation of OH and its detection by laser induced fluorescence (PLP-LIF) with in-situ, optical measurement of both NO_2 and H_2O we were able to show that (in contrast to previous investigations) the presence of H_2O increases the rate coefficient significantly. We derive a rate coefficient for H_2O bath gas at the low-pressure limit (k_0^{H2O}) of 15.9 × 10^{-30} cm^6 $molecule^{-2}$ s^{-1}. This indicates that H_2O is a more efficient collisional quencher (by a factor of ≈ 6) of the initially formed HO-NO_2 association complex than N_2 and a factor ≈ 8 more efficient than O_2. Ignoring the effect of water-vapour will lead to an underestimation of the rate coefficient by up to 15% e.g. in the tropical boundary layer. Combining the new experimental results from this study with those from the companion paper in which we report rate coefficients obtained in N_2 and O_2 bath gases (Amedro et al., 2019) we derive a new parameterisation for atmospheric modelling of the OH + NO_2 reaction and use this in a chemical transport model (EMAC) to examine the impact of the new data on the global distribution of NO_2, HNO_3 and OH. Use of the new parameters (rather than those given in the IUPAC and NASA evaluations) result in significant changes in the HNO_3 / NO_2 ratio and NOx concentrations, the sign of which depends on which evaluation is used as reference. The model predicts the presence of HOONO (formed along with HNO_3 in the title reaction) in concentrations similar to those of HO_2NO_2 at the tropical tropopause.

1 Introduction

In our recent study of the title reaction (Amedro et al., 2019), we reported extensive measurements of the rate constant (k_1) for the termolecular reaction between OH and NO_2 (R1) in N_2 and O_2 bath gas over a large range of temperature and pressures.

$$OH + NO_2 + M \rightarrow HNO_3 + M \quad \text{(R1a)}$$

$$\rightarrow HOONO + M \quad \text{(R1b)}$$

Reaction (R1) converts NO_2 to nitric acid (HNO_3) and peroxynitrous acid (HOONO), and its rate strongly influences the relative abundance of atmospheric NO_x (NO_2 + NO) and longer-lived "reservoirs" of NO_x which include e.g. HNO_3 and organic nitrates. It also converts OH (the main initiator of atmospheric oxidation) to a long-lived reservoir, HNO_3. As the abundance of OH and NO_x directly impact on photochemical ozone formation and the lifetimes of greenhouse gases, reaction (R1) may be considered one of the most important gas-phase processes in atmospheric science (Newsome and Evans, 2017). As outlined by Amedro et al. (2019), the rate coefficients and product-branching for this reaction are dependent on pressure and temperature and also on the bath-gas identity, i.e. the identity of the collision partner, M in reaction (R1). The per collision efficiency of energy transfer from the initially "hot" association complex to bath gas can vary considerably, with more complex bath gases possessing more degrees of freedom and bonds with similar vibrational frequencies to those in the association complex being generally more efficient. In this sense, we may expect H_2O to be better than N_2 or O_2 in quenching $[HO\text{-}NO_2]^{\#}$. In this second part of our study of the reaction between OH and NO_2, we extend the experiments to H_2O and He bath-gases. After N_2 (≈ 78%) and O_2 (≈ 21%) water vapour is the third most abundant gaseous species in the lower atmosphere. Its

concentration is highly variable in time and space, varying in mixing ratio from a few percent at sea level to parts-per-million in the stratosphere. Most of the atmosphere's water vapour is present in the planetary boundary layer where its average mixing ratio on the global scale is ≈ 1% but which may exceed 5% in tropical regions.

The effect of water vapour on gas-phase radical-reactions has been the subject of numerous studies (Buszek et al., 2011) and is sometimes interpreted in terms of formation of H_2O-radical complexes leading, via a chaperone type mechanism, to an increase in the rate constant. An important example of this is the HO_2 self-reaction for which the rate constant increases by a factor of up to two in the presence of water vapour due to formation of an HO_2-H_2O complex (Lii et al., 1981; Kircher and Sander, 1984). Theoretical calculations (Allodi et al., 2006; Sadanaga et al., 2006; Thomsen et al., 2012) suggest that, under our experimental conditions, the fraction of OH and NO_2 clustered with H_2O is < 0.1 % which is insufficient to significantly impact on k_1.

On the other hand, the role of H_2O as a collision partner in termolecular, atmospheric reactions has rarely been reported though its potential impact has been highlighted (Troe, 2003). Indeed, water vapour is known to be a more efficient third-body collider, by up to an order of magnitude compared to N_2 in termolecular reactions such as H + H + M, H + OH + M and H + O_2 + M (Getzinger and Blair, 1969; Michael et al., 2002; Fernandes et al., 2008; Shao et al., 2019).

The conclusions of three previous experiments examining the role of H_2O in kinetic studies of reaction (R1) are highly divergent, with the addition of H_2O found to 1) increase the rate coefficient (Simonaitis and Heicklen, 1972), 2) have no measureable effect (D'Ottone et al., 2001) or 3) even reduce it (Sadanaga et al., 2006). The overall aim of this research was to clarify these differences and provide quantitative data on the third-body efficiency of H_2O for the title reaction. Based on the kinetic data for the water vapour effect reported in this manuscript and in N_2 and O_2 presented in the first part of this study (Amedro et al., 2019) we have generated a new parameterisation for the overall rate coefficient, k_1, and examined its impact on atmospheric OH, NO_X and NO_Y in a global chemical transport model.

2. Experimental details

The details of the experimental set-up have been published previously (Wollenhaupt et al., 2000; Amedro et al., 2019) and only a brief description is given here.

2.1 PLP-LIF technique

The experiments were carried out in a quartz reactor of volume 500 cm^3 which was thermostatted to the desired temperature by circulating a 60:40 mixture of ethylene glycol-water. The pressure in the reactor was monitored with 100 and 1000 Torr capacitance manometers. Flow rates were chosen so that a fresh gas sample was available for photolysis at each laser pulse (laser frequency, 10 Hz), thus preventing build-up of products. Pulses of 248 nm laser light (≈ 20 ns) for OH generation from HNO_3, H_2O_2 and O_3/H_2O precursors were provided by an excimer laser (Compex 205 F, Coherent) operated using KrF.

$HNO_3 + h\nu$ (248 nm) → OH + NO_2 (R2)

$H_2O_2 + h\nu$ (248 nm) → 2 OH (R3)

$O_3 + h\nu$ (248 nm) → $O(^1D) + O_2$ (R4)

$O(^1D) + H_2O$ → 2 OH (R5)

OH concentrations (10^{11} – 10^{12} molecule cm^{-3}) were similar to those reported by Amedro et al. (2019) and the same arguments, which rule out significant influence of secondary reactions, apply. The concentration ranges of the H_2O_2, HNO_3 and O_3 precursors are listed in the notes to Tables 1 and 2.

OH was detected following excitation of the OH $A^2\Sigma(v'=1) \leftarrow X^2\Pi(v''=0)$ transition (Q11(1) at 281.997 nm using a YAG-pumped dye laser (Quantel-Brilliant B and Lambda-Physik Scanmate). OH fluorescence was detected by a photomultiplier tube (PMT) screened by a 309 nm interference filter and a BG 26 glass cut-off filter.

2.2 On-line absorption measurement of NO_2 and H_2O concentration

As discussed by Amedro et al. (2019), the determination of the NO_2 concentration is critical for accurate measurement of k_1. We therefore deployed in-situ, broad-band (405 – 440 nm) and single wavelength (365 nm) optical absorption spectroscopy. The former was located prior (in flow) to the quartz-reactor, the latter was located behind the quartz-reactor. Using the broadband cell, the NO_2 concentration was retrieved by least square fitting from 405 to 440 nm to a reference spectrum (Vandaele et al., 2002) degraded to the resolution of our spectrometer. Simultaneously, we measured NO_2 at 365 nm using the absorption cross-section 5.89×10^{-19} cm^2 $molecule^{-1}$ determined previously by Amedro et al. (2019) who give a detailed description of the NO_2 concentration measurements and the choice of reference spectrum. For the temperatures used in this study, corrections to the NO_2 concentration due to formation of the N_2O_4 dimer were not necessary.

For the present experiments, a third absorption cell (l = 40 cm) was placed downstream of the quartz-reactor to measure the H_2O concentration at 184.95 nm. This set-up used a low-pressure Hg-Penray lamp isolated with a 185 nm interference filter as light source. Optical extinction was converted to concentrations using a cross-section of 7.14×10^{-20} cm^2 $molecule^{-1}$ (Cantrell et al., 1997).

2.3 Chemicals

N_2 and He (Westfalen 99.999%) were used without further purification. H_2O_2 (AppliChem, 50 wt. %) was concentrated to >90 wt.% by vacuum distillation. Anhydrous nitric acid was prepared by mixing KNO_3 (Sigma Aldrich, 99%) and H_2SO_4 (Roth, 98%), and condensing HNO_3 vapour into a liquid nitrogen trap. NO (3.5 AirLiquide) was purified of other nitrogen oxides by fractional, vacuum distillation and then converted to NO_2 via reaction with a large excess of O_2. The NO_2 thus made was trapped in liquid N_2 and the excess O_2 was pumped out. The resulting NO_2 was stored as a mixture of ~0.5% NO_2 in N_2 or ~5.5% NO_2 in He. Distilled H_2O (Merck, Liquid Chromatography grade) was degassed before use and kept at constant temperature.

3 Results and Discussion

3.1 Measurements of k_1 in He bath-gas and comparison with literature

Our study of the role of H_2O as collision partner in reaction (R1) was carried out in mixtures of He-H_2O and N_2-H_2O. In order to separate the effects of H_2O and He, we also required accurate rate coefficients for pure He bath gas, which we describe below. As for the N_2 and O_2 bath-gas datasets (Amedro et al., 2019), the experiments were carried out under pseudo-first-order conditions ($[NO_2] >> [OH]$) so that Eqn. 1-2 describe the decay of OH and the derivation of the bimolecular rate coefficient, k_1.

$$[OH]_t = [OH]_0 \exp(-k' t) \quad (1)$$

where $[OH]_t$ is the concentration (molecule cm^{-3}) at time t after the laser pulse. k' is the pseudo-first order rate coefficient and is defined as

$$k' = k_1[NO_2] + k_d \quad (2)$$

where k_d (s^{-1}) accounts for OH-loss due to diffusion out of the reaction zone and reaction with its photolytic precursors such as HNO_3 or H_2O_2.

An exemplary dataset illustrating OH decays and a plot of k' versus [NO_2] is given in Fig. S1 of the supplementary information).

Values of k_1 obtained in He bath-gas (25-690 Torr, 292 K) are summarised in Fig. 1 and 2 and listed in Table 1. The kinetics of termolecular reactions can be described by the Lindemann-Hinshelwood mechanism whereby the rate constant at the low-pressure limit (k_0, units in cm^6 $molecule^{-2}$ s^{-1}) is proportional to the pressure and at the high pressure limit (k_∞, units in cm^3 $molecule^{-1}$ s^{-1}) is independent of pressure. In the intermediate pressure range, the fall-off regime, the rate coefficient is a function of both low-pressure (k_0) and high-pressure (k_∞) rate coefficients and the (reaction-partner dependent) broadening factor F which accounts for the lower rate constant measured in the fall-off regime than predicted by the Lindemann-Hinshelwood mechanism reactions (Troe, 1983). Under the conditions of T and p relevant for atmospheric chemistry, the title reaction is in the fall-off regime.

$$k = \frac{k_0[\mathrm{M}]k_\infty}{k_0[\mathrm{M}] + k_\infty} F \quad (3)$$

The solid lines in Figs. 1 and 2 are fits according to the Troe formalism for termolecular reactions (Troe, 1983) as adopted by the IUPAC panel in their evaluation of atmospheric reactions:

$$k(P,T) = \frac{k_0\left(\frac{T}{300}\right)^{-m}[\mathrm{M}]k_\infty\left(\frac{T}{300}\right)^{-n}}{k_0\left(\frac{T}{300}\right)^{-m}[\mathrm{M}] + k_\infty\left(\frac{T}{300}\right)^{-n}} F \quad (4)$$

where T is the temperature in Kelvin, [M] is the bath-gas concentration in molecule cm^{-3}, m and n are dimensionless temperature exponents.

The broadening factor, F, is:

$$\log F = \frac{log F_c}{1 + \left[log\left(\frac{k_0\left(\frac{T}{300}\right)^{-m}[\mathrm{M}]}{k_\infty\left(\frac{T}{300}\right)^{-n}} \right) / N \right]^2} \quad (5)$$

Where $N = [0.75 - 1.27 \log F_c]$ and F_c is the broadening factor at the centre of the fall-off curve.

As discussed in some detail in the first part of our studies of the title reaction (Amedro et al., 2019), the low- or high-pressure rate constants for the title reaction (k_0 and k_∞) are not well defined by existing data sets, which do not deliver sufficiently accurate rate coefficient at very low pressures (< 1 mbar) or at very high pressures (> 500 bar). Studies in which k_∞ has been derived from rates of vibrational relaxation of OH (Smith and Williams, 1985; D'Ottone et al., 2005), return values of k_∞ that provide some constraint on its value, but the associated uncertainty is too large to consider this parameter well defined.

In our first paper, Amedro et al. (2019) describe highly accurate measurements of k_1 over a wide range temperatures and pressures in the fall-off regime. From measurements of k_1 in N_2 bath-gas, we retrieved values of k_0 and k_∞ of 2.6×10^{-30} cm^6 $molecule^{-2}$ s^{-1} and 6.3×10^{-11} cm^3 $molecule^{-1}$ s^{-1}, respectively, by fixing F_c to a value of 0.39 which has theoretical basis (Cobos and Troe, 2003). The reasons for choosing this value of F_c are discussed in Amedro et al. (2019). Note that whereas k_0 is dependent on the bath-gas used, at the high-pressure limit, k_∞ should be the same in N_2, O_2, He or H_2O bath gases.

In Fig.1 we display pressure dependent rate coefficients (solid, black squares) obtained in He bath-gas at 292 K. The black line is a fit (Eqn. 4) to our data with k_∞ fixed to 6.3×10^{-11} cm^3 $molecule^{-1}$ s^{-1} and $n = 0$ as derived from an extensive dataset obtained using N_2 bath-gas (Amedro et al., 2019). For this dataset, the best fit is obtained with $F_c = 0.32$, and $k_0^{He} = 1.4 \times 10^{-30}$ cm^6 $molecule^{-2}$ s^{-1}. When using $F_c = 0.39$ (i.e. same value as that obtained in N_2 bath-gas), the fit slightly overestimates (~5 %) the measurements at pressures above ~300 Torr whereas it underestimates by 10 % at lower pressures (Fig. S2). We note that using a higher $F_c = 0.39$ resulted in a lower value of k_0^{He} equal to 1.0×10^{-30} cm^6 $molecule^{-2}$ s^{-1}. The T-dependence factor in He, m(He), was determined to be 3.1 over the temperature range from 277 to 332 K (Table 1 and Figure S6).

The high precision of our measurements in He and N_2 indicates that different broadening factors (F_c) are required to interpret the pressure dependence of k_1 obtained in N_2 and He. This can be rationalized by considering that F_c is the product of strong-collision (F_c^{SC}) and weak-collision (F_c^{WC}) components (Eqn. 6-8) (Gilbert et al., 1983; Troe, 1983; Troe and Ushakov, 2011)

$$F_c \approx F_c^{SC} F_c^{WC} \quad (6)$$

$$F_c^{SC} \approx {S_K}^{-0.62} \approx \left(1 + \frac{r}{2}\right)^{-0.62} \quad (7)$$

$$F_c^{WC} \approx {\beta_c}^{0.14} \quad (8)$$

Here, S_K is the Kassel parameter and r is the total number of external rotational modes of the reactants (equal to 5 in the reaction between OH and NO_2) and β_c is the collision efficiency. While the strong collision component is independent of bath gas ($F_c^{SC} \approx 0.46$ for the title reaction) a change in F_c^{WC} due to a lower collision efficiency (β_c) of He relative to N_2 is likely. The collision efficiency for N_2 which was used to calculate $F_c = 0.39$ was $\beta_c(N_2) \approx 0.3$ (Troe, 2001). The value of $F_c = 0.32$ from our He data implies $\beta_c(\mathrm{He}) \approx 0.08$, a factor 3.7 times lower than $\beta_c(N_2)$. A large difference in collision efficiency between N_2 and He is consistent with theoretical calculations (Glänzer and Troe, 1974; Troe, 2001; Golden et al., 2003).

In Fig 1, we also compare our measurements of k_1 in He with data collected in the same pressure range using similar techniques. The three first measurements (Morley and Smith, 1972; Anastasi and Smith, 1976; Wine et al., 1979) used flash photolysis of H_2O as a OH precursor with detection of OH by resonance fluorescence. Morley and Smith (1972) reported rate coefficients at pressures of 20 to 280 Torr at room temperature with the NO_2 concentration calculated manometrically. Our parametrisation agrees within the combined uncertainty of both measurements (Figure S3). Anastasi and Smith (1976) reported one value of k_1 at 25 Torr He which is ≈ 20 % lower than our measurement. Wine et al. (1979) presented values of k_1 at 3 pressures of He. The agreement with our parameterisation at the lowest two pressures is excellent but a deviation of ≈ 20% is observed at the highest pressure (Figure S4). As both studies measured NO_2 concentrations using optical absorption at 365 nm, the ≈ 20% difference is significant. Most recently, D'Ottone et al. (2001) reported rate coefficients from 30 to 600 Torr He using a very similar approach to ours i.e. PLP-LIF technique with in situ measurements of NO_2 by absorption at 365 nm. The disagreement (up to 40%) between our measurements and theirs exceed the combined reported uncertainty (Figure S5). While it is unclear what could have caused the discrepancy, we note that the data of D'Ottone et al. (2001) are significantly more scattered and do not describe a smooth increase in rate coefficient with pressure as expected from termolecular reactions in the fall-off regime. This would appear to indicate an underestimation of the total uncertainty in their study.

Figure 2 extends the pressure range to additionally display data obtained in low pressure flow-tubes (Westenberg and Dehaas, 1972; Anderson et al., 1974; Erler et al., 1977; Anderson, 1980) and the high-pressure measurements by Hippler et al. (2006). At low pressures our data is in excellent agreement (within 10%) with the data of Erler et al. (1977) but predict values ≈ 40% lower than those reported by Westenberg and Dehaas (1972) and Anderson (1980). The data of Anderson et al. (1974) display a large intercept (4.9×10^{-14} cm^3 molecule^{-1} s^{-1}) at zero pressure, which is attributed to a second-order heterogeneous removal rate constant. As indicated in a critical assessment of the low-pressure data by Amedro et al. (2019) it is unclear whether one can simply subtract a constant value equal to the intercept (obtained from a linear fit) to each data point. If we were to do so, the work by Anderson et al. (1974) would be in very good agreement with the low pressure study by Erler et al. (1977) as well as with our parametrisation extended to low pressures. Additionally, Amedro et al. (2019) demonstrated that, owing to the large, asymmetric broadening of fall-off for this reaction the assumption that the rate coefficient is in the low pressure limit at N_2 pressures of 0.5 Torr < p < 10 Torr is invalid and leads to underestimation of k_0. This observation is still true of datasets obtained at low pressures of He, so that while very good agreement is observed between our parametrisation and individual rate coefficients obtained between 3 and 8 Torr of He, reported values of k_0^{He} are 40 % lower than our values obtained from the fall-off analysis. As indicated in Fig. 2, our parametrisation of k_1 in He is in very good agreement with the high pressure data reported by Hippler et al. (2006).

3.2 Influence of H_2O on k_1

As mentioned above, the effect of water vapour on k_1 was determined in mixtures of H_2O with both N_2 and He. This is because the vapour pressure of H_2O at room temperature (≈ 17 Torr at 293 K) is too low to enable experiments in pure H_2O bath gas to be conducted using our instrument. The measurements were performed at low density ([M] = 1.6×10^{18} molecule cm^{-3}; 50 Torr at 293 K) where the relative increase of k_1 in the presence of H_2O is pronounced, resulting in greater accuracy in the determination of k_0^{H2O}. Experimental data on the influence of H_2O on k_1 was obtained in N_2-H_2O and He-H_2O mixtures by varying the H_2O mixing ratio, x_{H2O}, from 0.05 to 0.27 ([H_2O] = 0.9 - 4.5 × 10^{17} molecule cm^{-3}) while keeping the total pressure constant at 50 Torr. Under these conditions, the addition of H_2O resulted in an increase in k_1 up to a factor of two as illustrated by the datasets of Fig. 3 in which the increase in slope as more water-vapour is added is proportional to the increase in k_1 (Eqn. 2). At the highest concentration of water vapour (4.5×10^{17} molecule cm^{-3}) the rate coefficient in He-H_2O increased by > factor 3 compared to that obtained in pure He (see Table 1).

In order to determine the temperature dependence of the enhancement in k_1 caused by the presence of water, the experiments in He were carried out at 3 different temperatures (277, 291 and 332 K). The values of k_1 obtained from these experiments are plotted versus the mole-fraction of H_2O in Fig. 4b. At the pressures used in our experiments, k_1 displays fall-off, precluding direct measurement of k_0^{H2O}.

~~In other words,~~The total rate constant measured in e.g. a H_2O-N_2 bath gas is not equal to the sum of the individual rate constants calculated from the mixing ratios of N_2 and H_2O i.e. $k_{N2\text{-}H2O} \neq k_{N2} + k_{H2O}$. $k_{N2\text{-}H2O}$ is only equal to the sum of k_{N2} and k_{H2O} at the low pressure limit (<< 1 Torr in the case of the OH reaction with NO_2) and under certain conditions where gas mixtures are composed of strong colliders and/or have similar collision efficiencies (Troe, 1980; Burke and Song, 2017). Additionally, at the high-pressure end of the fall-off curve, the rate coefficient is independent of bath gas composition. To be able to make a reasonable prediction of this effect under atmospheric conditions where the mole fraction of water vapour, x_{H2O}, can be as large as 0.05, we analysed our measurements using two different approaches to determine k_0^{H2O}. In the first case, the low pressure rate constant in a N_2-H_2O mixture is defined as the sum of two individual low pressure limit rate constants,

$$k(P,T) = \frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]k_\infty\left(\frac{T}{300}\right)^{-n}}{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M] + k_\infty\left(\frac{T}{300}\right)^{-n}} F \quad (9)$$

where x_{N2} and x_{H2O} are the mixing ratio for N_2 and H_2O respectively, k_0^{N2} and k_0^{H2O} are low-pressure limiting rate constants (units of cm^6 $molecule^{-2}$ s^{-1}) for pure N_2 and H_2O, k_∞ is the high-pressure limit rate constant (units of cm^3 $molecule^{-1}$ s^{-1}), T is the temperature in Kelvin, [M] is the molecular density (molecule cm^{-3}) and m, n and o are dimensionless temperature exponents.

The broadening factor, F, is:

$$\log F = \frac{log F_c}{1 + \left[log\left(\frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]}{k_\infty\left(\frac{T}{300}\right)^{-n}}\right)/N\right]^2} \quad (10)$$

Where N = [0.75 - 1.27 log F_c] and F_c is the broadening factor at the centre of the fall-off curve.

In the second approach, we follow Burke and Song (2017) where, additionally to the low pressure limiting rate coefficients, the broadening factors for each bath gas are also mixed linearly and log $F^{N2\text{-}H2O}$ is defined as

$$\log F^{N2\text{-}H2O} = \tilde{X}_{N2} \log F^{N2} + \tilde{X}_{H2O} \log F^{H2O} \quad (11)$$

$$\text{where} \quad \tilde{X}_{N2} = \frac{x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m}[M]}{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]}; \quad \tilde{X}_{H2O} = \frac{x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-m}[M]}{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]} \quad (12)$$

$$\log F^{N2} = \frac{log F_c^{N2}}{1+\left[log\left(\frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m}+x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]}{k_\infty\left(\frac{T}{300}\right)^{-n}}\right)/(0.75-1.27\log F_c^{N2})\right]^2} \quad (13)$$

$$\log F^{H2O} = \frac{log F_c^{H2O}}{1+\left[log\left(\frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m}+x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)[M]}{k_\infty\left(\frac{T}{300}\right)^{-n}}\right)/(0.75-1.27\log F_c^{H2O})\right]^2} \quad (14)$$

where F_c^{N2} and F_c^{H2O} are the broadening factor at the centre of the fall off curve for N_2 and H_2O.

In the case where two bath gases have identical (or very similar) values of F_c, the two approaches result in identical predictions and the first approach will be preferred for its simplicity. This is the case for N_2 and H_2O bath gases. However, when two bath-gases have significantly different values of F_c (as is the case for He-H_2O mixtures, see below) the second approach provides a more accurate parameterisation.

3.2.1 Parameterisation of k_1 from data obtained in N_2-H_2O and He-H_2O bath gases

Values of k_1 obtained in N_2–H_2O and He-H_2O bath gases are listed in Table 2. Each rate coefficient obtained in N_2-H_2O bath gas was defined by 5 parameters: the mixing ratio of N_2 and H_2O (x_{N2} and x_{H2O}) the overall rate coefficient (k_1) the molecular density [M] and the temperature T. We performed a multivariate fit of the N_2-H_2O dataset with k_0^{H2O} as variable, all other parameters fixed with: $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-1} s^{-1}, $k_0^{N2} = 2.6 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1} and $m = 3.6$ as derived in Amedro et al. (2019), o was fixed to 3.4 (see below) and F_c was held at 0.39 making the assumption that the broadening factors at the centre of the fall-off curve for H_2O and N_2 were identical. The fit to the data returned $k_0^{H2O} = (15.9 \pm 0.7) \times 10^{-30}$ cm^6 molecule^{-2} s^{-1} where the uncertainty is 2σ (statistical only). The solid black line on the upper panel of Fig. 4a represents the parametrisation for a varying fraction of H_2O in N_2 at a total pressure of 50 Torr using the parameters given above. Equating F_c^{H2O} and F_c^{N2} simplifies the analysis, though it is likely that $F_c^{H2O} > F_c^{N2}$ as the collision efficiency (β_c) is likely to be larger for H_2O than for N_2. We found that the He-H_2O data cannot be modelled assuming the same F_c for both He and H_2O bath gas and the approach of Burke and Song (2017) was therefore preferred. In order to analyse the data we fixed the following parameters: $k_0^{H2O} = 15.9 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $F_c^{H2O} = 0.39$, $F_c^{He} = 0.32$ and $k_0^{He} = 1.4 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1} and $m = 3.1$ to derive $o = (3.4 \pm 0.8)$ (2σ, statistical only), which describes the temperature dependence of the low pressure limit in H_2O as depicted in Fig. 4b.

There is clearly some uncertainty related to the arbitrary use of $F_c^{H2O} = 0.39$. For example, if we were to use analyse the data in N_2-H_2O using $F_c^{H2O} = 0.6$ and the linear mixing method we retrieve $k_0^{H2O} = 10 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, which is ≈ 50% lower than our preferred value. The effect of the different analyses can be assessed by comparing the predicted impact of H_2O on k_1 at 80% relative humidity, 1000 mbar and 313 K. If we set $F_c^{H2O} = 0.39$ we predict that the effect of H_2O is to increase k_1 by 15% while choosing $F_c^{H2O} = 0.6$ results in an increase of 20%. Theoretical calculation of the relative values of Fc in N_2, O_2 and H_2O bath gases input would be useful to reduce this uncertainty. Our data indicate a significant, positive trend in k_1 when adding H_2O. As discussed above, more efficient energy transfer from [HO-NO_2]$^\#$ in collision with H_2O compared to N_2 is intuitive and supported by the present dataset as well as that of Simonaitis and Heicklen (1972) who derived $k_0^{H2O} = 11 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}. Given the complexity of the analysis, this may be considered to be in good agreement. This result is however not consistent with the observations of D'Ottone et al. (2001) who report no significant change in k_1 in 150 Torr of He when adding either 10 or 20 Torr of H_2O and is completely at odds with the conclusions of Sadanaga et al. (2006), who report a reduction in k_1 (by 18%) when adding 29.1 mbar of H_2O at atmospheric pressure. If our value for k_0^{H2O} is correct, D'Ottone et al. (2001) should have seen an increase in k_1 of ≈ 55% and Sadanaga et al. (2006) should have observed an increase of ≈ 5%.

A potential explanation for the very divergent observations of the effect of H_2O is the heterogeneous loss of NO_2 when adding H_2O. We tested for NO_2 loss in a set of experiments in which NO_2 and H_2O were monitored simultaneously while systematically varying the amount of H_2O. Our results indicated a reduction in the concentration of NO_2 by up to ≈ 20% as we increased the concentration of H_2O up to 4.5×10^{17} molecule cm^{-3}. Unless NO_2 is monitored in-situ (as in our experiments), 20% loss of NO_2 would lead to a similar size reduction in the OH decay constant and thus an underestimation of the rate coefficient. A fractional loss of NO_2 of this magnitude would explain why Sadanaga et al. (2006) found an apparent reduction in k_1 when adding H_2O.

However, the situation becomes more complex if NO_2 is converted to trace gases that are reactive towards OH. For this reason, we performed an additional experiment to investigate whether NO_2 was converted via reaction with H_2O on surfaces to HONO and/or HNO_3. Note that conversion of NO_2 to HONO at low pressures (e.g. 50 Torr) would result in an increase in the OH decay constant ($k_{OH+HONO} > k_{OH+NO2}$), whereas conversion of NO_2 to HNO_3 would result in a decrease ($k_{OH+HNO3} < k_{OH+NO2}$).

In order to test for the presence of HONO, we modified the broadband absorption set-up by replacing the halogen lamp with a deuterium lamp, allowing us to detect HONO around 350 nm as well as NO_2. The optical absorption of NO_2 and HONO (340 – 380 nm) was monitored in a flow of NO_2 (1.7×10^{15} cm^{-3}) at 50 Torr He in the absence and presence of H_2O ($[H_2O] = 4.5 \times 10^{17}$ molecule cm^{-3}, the maximum concentration used in this work). A depletion in NO_2 of 21% (3.7×10^{14} molecule cm^{-3}) was observed when H_2O was added. An analysis of the spectra with and without H_2O (Fig. S7) enabled us to establish an upper limit to the HONO concentration of ≈ 1×10^{13} molecule cm^{-3}, which would corresponds to just 3% of the NO_2 lost. At this concentration, HONO does not significantly increase the loss-rate of OH (< 3% using a rate coefficient for reaction of OH with HONO of 6.0×10^{-12} cm^3 $molecule^{-1}$ s^{-1} (IUPAC, 2019)). In the same experiment, we also recorded the optical density at 185 nm where H_2O, NO_2 and HNO_3 all absorb. Despite the large HNO_3 absorption cross-section at this wavelength (1.6×10^{-17} cm^2 $molecule^{-1}$, Dulitz et al. (2018)) we found no evidence for HNO_3 formation, indicating that the NO_2 lost was not converted to gas-phase HNO_3. Given its great affinity for glass in the presence of H_2O, we expect that any HNO_3 formed is strongly partitioned to the walls of the reactor. The tests indicate that, on the time scales of our experiments, NO_2 is lost irreversibly on the humidified walls of our experiment. The maximum concentration of H_2O used in this experiment, 4.5×10^{17} molecule cm^{-3}, corresponding to a relative humidity of 80% (at 292 K) so that H_2O condensation is not expected.

It is difficult to establish whether our observations of significant NO_2 loss can explain the result of D'Ottone et al. (2001), who did not observe an enhancement in k_1. D'Ottone et al. (2001) did not state whether, in their experiments, NO_2 and H_2O were monitored simultaneously. Also, our observed loss of NO_2 is not necessarily transferable to other studies as the heterogeneous loss of NO_2 will vary from one experimental set-up to the next, as residence times and surface areas may vary substantially.

A very simple calculation serves to illustrate the role of water vapour as a third-body quencher for the title reaction. We consider e.g. the tropical boundary layer with a temperature of 30 °C and a relative humidity of 80% at a total pressure of 1 bar. The pressure of water vapour is 34 mbar, those of O_2 and N_2 are then 210 and 756 mbar, respectively. A rough contribution of each quenching gas to the overall rate coefficient can be calculated from the respective low-pressure rate coefficients. For N_2, O_2 and H_2O these are (in units of 10^{-30} cm^3 $molecule^{-1}$ s^{-1}) 2.6, 2.0 and 15.9. Water vapour is therefore a factor ≈ 8 more efficient than O_2, and a factor ≈ 6 more efficient than N_2 as a quencher of the HO-NO_2 intermediate, which is qualitatively consistent with known strong binding (40 kJ mol^{-1}) in the HNO_3–H_2O complex (Tao et al., 1996).

For our tropical boundary layer case-study, in which the O_2 pressure is only a factor of six greater than that of H_2O, we calculate that H_2O contributes more to the rate coefficient of the title reaction than does O_2. Clearly, the neglect of including the quenching effect of H_2O leads to underestimation (in the boundary layer) of the rate coefficient for this centrally important atmospheric reaction.

In order to assess both the effect of H_2O (this work) and the new parameterisation for k_1 in N_2 and O_2 bath-gases presented in first part of this study (Amedro et al., 2019), we have used a 3D chemical transport model (EMAC, see below) to explore the impact on a global scale.

3.3. Atmospheric modelling of the OH + NO_2 reaction including the effect of water vapour

The EMAC (ECHAM-MESSy Atmospheric Chemistry) model employed is a numerical chemistry and climate simulation system (Jöckel et al., 2006; Jöckel et al., 2010) using the 5th generation European Centre Hamburg general circulation model (ECHAM5, Roeckner et al. (2006)) as core atmospheric general circulation model. For the present study we applied EMAC (ECHAM5 version 5.3.02, MESSy version 2.53.0) in the T42L47MA-resolution, i.e. with a spherical truncation of T42 (corresponding to a quadratic Gaussian grid of approx. 2.8 by 2.8 degrees in latitude and longitude) with 47 vertical hybrid pressure levels up to 0.01 hPa. The model has been weakly nudged in spectral space, nudging temperature, vorticity, divergence and surface pressure (Jeuken et al., 1996). The chemical mechanism scheme adopted (MOM, Mainz Organic Mechanism) includes oxidation of isoprene, saturated and unsaturated hydrocarbons, including terpenes and aromatics (Sander et al., 2019). Further, tracer emissions and model set-up are similar to the one presented in Lelieveld et al. (2016a). EMAC model predictions have been evaluated against observations on several occasions (Pozzer et al., 2010; de Meij et al., 2012; Elshorbany et al., 2014; Yoon and Pozzer, 2014): For additional references, see http://www.messy-interface.org. For this study, EMAC was used in a chemical-transport model (CTM mode) (Deckert et al., 2011), i.e., by disabling feedbacks from photochemistry on radiation and dynamics. Two years were simulated (2009-2010), with the first year used as spin-up time.

The following parameterisation of k_1 was implemented in EMAC; values of each parameter are listed in Table 3.

$$k_1(P,T) = \frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{O2}k_0^{O2}\left(\frac{T}{300}\right)^{-q} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)Mk_\infty\left(\frac{T}{300}\right)^{-n}}{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{O2}k_0^{O2}\left(\frac{T}{300}\right)^{-q} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)M + k_\infty\left(\frac{T}{300}\right)^{-n}}F \quad (15)$$

The broadening factor, log F, is:

$$\log F = \frac{log F_c}{1+\left[log\left(\frac{\left(x_{N2}k_0^{N2}\left(\frac{T}{300}\right)^{-m} + x_{O2}k_0^{O2}\left(\frac{T}{300}\right)^{-q} + x_{H2O}k_0^{H2O}\left(\frac{T}{300}\right)^{-o}\right)M}{k_\infty\left(\frac{T}{300}\right)^{-n}}\right)/[0.75 - 1.27\log Fc]\right]^2} \quad (16)$$

As described in Section 1, the reaction between OH and NO_2 forms not only HNO_3 but also HOONO. HOONO decomposes rapidly at typical boundary layer temperatures but is long lived with respect to thermal dissociation at the temperatures found in the upper troposphere and lower stratosphere (UTLS).

HOONO + M $\rightarrow$ OH + NO_2 (R6)

The rate constant (k_6) for thermal decomposition of HOONO was calculated from the channel specific rate coefficient for its formation ($k_1\alpha$) and an equilibrium coefficient: $k_6 = k_1\alpha / K_{eq}$, where $K_{eq} = 3.5 \times 10^{-27}\exp(10135/T)$ (Burkholder et al., 2015; IUPAC, 2019) based on the analysis of (Golden et al., 2003). The branching ratio to HOONO formation (α) was adapted from the present IUPAC recommendations for k_{1a} and k_{1b} which were derived from experimental work (Hippler et al., 2006; Mollner et al., 2010) and theoretical analysis (Troe, 2012). The IUPAC recommendations were augmented with a pressure independent HOONO yield of 0.035 to better represent the dataset of Mollner et al. (2010) who detected HOONO directly at room temperature. We assume α is independent of water vapour. The expression used and a plot of α at different temperatures and pressures is given in Fig. S8 of the supplementary information.

In the absence of experimental data on the reactions of HOONO with OH or on its photolysis, we follow the approach of Golden and Smith (2000) and set these equal to those for HO_2NO_2 :

HOONO + OH $\rightarrow$ H_2O + NO_3 (R7)

HOONO + $h\nu$ $\rightarrow$ HO + NO_2 (R8)

In Fig. 5, we illustrate the global impact (annual average) of H_2O-vapour on the rate coefficient. We plot the fractional reduction in k_1 at the Earth's surface when setting x_{H2O} to zero rather than using the EMAC global water-vapour fields. We focus on the boundary layer as the H_2O concentration is largest here and decreases rapidly with altitude.

As expected, the greatest effect is found in warm, tropical regions where neglecting the impact of water vapour results in an average underestimation of the rate coefficient by up to ≈ 8 %. At higher/lower latitudes the effect is diminished and water vapour accounts for only 3-4 % of the overall rate coefficient at 40 ° N/S. The presence of water vapour does not impact on values of k_1 above the boundary layer.

Our experimental data do not give insight into whether the H_2O-induced enhancement in k_1 is accompanied by a change in the branching ratio to favour either HNO_3 or HOONO. However, as the formation of HOONO is favoured at high pressures (more effective collisional deactivation) it is possible that the HOONO yield may be enhanced relative to HNO_3 in the presence of H_2O. If this is the case, the increase in rate coefficient at high water vapour levels (e.g. in the tropical lower troposphere) may be to some extent offset by the subsequent thermal dissociation of HOONO in these warm regions.

As described by Amedro et al. (2019) (Fig. 1 of their manuscript) two expert panels (IUPAC, NASA) evaluating kinetic data for use in atmospheric modelling fail to reach consensus for the title reaction, with the preferred rate coefficients differing by as much as 50% in the cold UTLS. For this reason we have calculated values of $\frac{k_1^{\text{NASA}}}{k_1^{\text{this work}}}$ and $\frac{k_1^{\text{IUPAC}}}{k_1^{\text{this work}}}$ at different altitudes and latitudes (i.e. at different temperatures and pressures). We parameterized the rate coefficient using the expressions given in this work (Eqn. 15, Table 3) and in the latest evaluations of IUPAC (k_1 last evaluated in 2017 (IUPAC, 2019)) and NASA (last evaluation published in 2015 (Burkholder et al., 2015)). As displayed in Fig. 6, values of $\frac{k_1^{\text{NASA}}}{k_1^{\text{this work}}}$ and $\frac{k_1^{\text{IUPAC}}}{k_1^{\text{this work}}}$ vary greatly with pressure and temperature and thus altitude. The NASA recommendations are always slightly lower but in good agreement (≤ 10%) for most of the troposphere, with larger differences ($(\frac{k_1^{\text{NASA}}}{k_1^{\text{this work}}})$ always < 1) only observed in the lower and mid-stratosphere. At altitudes above ≈ 30 km the ratio decreases to ≈ 0.8. A comparison with the rate coefficient derived from the IUPAC parameterization, shows that $\frac{k_1^{\text{IUPAC}}}{k_1^{\text{this work}}}$ varies from ≈ 0.9 at the surface to ≈ 1.1 at the tropopause but increases to >1.3 at the low pressures and temperatures that reign at 30 km and above. At high altitudes (low pressure and temperature) the rate coefficients that the evaluation panels recommend are strongly biased by choice of the rate coefficient (and its temperature dependence) at the low pressure limit. As discussed by Amedro et al. (2019) the available experimental data at low pressures and temperature are not of sufficient accuracy to use as basis for recommendation of k_0 and this is reflected in the highly divergent values of k_1 under these conditions.

As mentioned above, the atmospheric HNO_3 / NO_2 ratio is expected to be highly sensitive to the rate coefficient k_1, with an increase in k_1 resulting in an increase in the HNO_3 / NO_2 ratio and vice versa. The HNO_3 / NO_2 ratio also depends on the concentration of OH and thus the effect of using different values of k_1 will be most apparent in regions where the greatest OH concentrations are found, i.e. at low latitudes. At higher latitudes, especially in winter months where solar insolation is weak and OH levels are relatively low, the HNO_3 / NO_2 ratio will also be impacted by nighttime conversion of NO_2 to N_2O_5 and finally, via heterogeneous hydrolysis, to HNO_3. In Fig. 7 we plot zonally and yearly averaged model values of $\frac{HNO3}{NO2}\,(\text{IUPAC})/\frac{HNO3}{NO2}\,(\text{this work})$ in the upper panel and $\frac{HNO3}{NO2}\,(\text{NASA})/\frac{HNO3}{NO2}\,(\text{this work})$ in the lower panel. Compared to the present parameterization of k_1, the IUPAC evaluation returns HNO_3 / NO_2 ratios that are between 0.9 and 1 throughout most of the lower and free-troposphere (up to ≈ 5 km) and larger HNO_3 / NO_2 ratios (factor of 1.1 to 1.15) above ≈ 10 km especially at the tropical tropopause. The divergence between the HNO_3 / NO_2 ratios increases as we move further into the stratosphere with $\frac{HNO3}{NO2}\,(\text{IUPAC})/\frac{HNO3}{NO2}\,(\text{this work})$ as large as 1.2 to 1.3 above 25 km. At the same time, NO_X levels (NO_X = NO + NO_2) decrease by a factor ≈ 0.95 (see Fig. S9 of the supplementary information). When we compare our parameterization with that of the NASA panel, the picture is largely reversed (lower panel). Again, we find reasonable agreement in the HNO_3 / NO_2 ratio in the lowermost atmosphere, but in this case lower values (0.8 to 0.9) in the lower stratosphere which are accompanied by a factor 1.06 change in NO_X concentrations (Fig. S9). For both the NASA and IUPAC parameterizations, the largest differences in the HNO_3 / NO_2 ratio compared to the present study are found higher in the atmosphere. The modelling studies confirm the

simple calculation of Amedro et al. (2019) (see Fig. 1 of their paper), showing that the IUPAC and NASA parameterizations result in very different values of k_1 in some parts of the atmosphere and will result in divergent predictions of partitioning of reactive nitrogen between NO_X and NO_Y. Use of the parameterization based on the present dataset lies roughly between the two evaluations, with best agreement observed with NASA for the lower atmosphere. However, as previous laboratory studies had not identified the important role of H_2O in the title reaction, which could therefore not be incorporated in either of the previous parameterizations, any agreement at better than 10% level is fortuitous, reflecting random cancelling of systematic bias.

As reaction with OH is the predominant sink for most atmospheric trace-gases, its concentration largely defines the oxidizing power of the atmosphere (Lelieveld et al., 2004; Lelieveld et al., 2008; Lelieveld et al., 2016b) and even changes of a few percent in its concentration are significant. An increase in the rate coefficient of the title reaction will reduce the atmospheric abundance of this centrally important radical. In Fig. S10 we illustrate the impact of using the parameterization of k_1 from the present study compared to the IUPAC and NASA recommendations. The upper panel in Fig. S10 plots the ratio of OH concentrations obtained when using the IUPAC parameterization and that from the present study, OH(IUPAC) / OH(this work). Throughout the troposphere OH(IUPAC) / OH(this work) deviates by only a few percent, with a value of 1.02 at the surface and 0.96 at the tropical tropopause. OH(NASA) / OH(this work) is also 1.02 at the surface but increases to 1.04 at the tropical tropopause as the NASA-derived value of k_1 is lower at the temperatures and pressures encountered in this part of the atmosphere. The weak effect of changing k_1 on OH at the surface reflects the fact that many reactions apart from that with NO_2 contribute to the overall sink term for OH in the lower troposphere.

Although our experiments do not give insight into the branching between formation of HOONO and HNO_3 in the title reaction, previous work predicts a significant yield of HOONO especially at low temperatures (see Fig S8). As the lifetime of HOONO with respect to re-dissociation to reactants is short at e.g. boundary layer temperatures (≈ 1s at 298 K and 1 bar pressure), its formation may be seen as an effective reduction in the rate coefficient for OH + NO_2 (Golden and Smith, 2000). However, its lifetime increases to several days at temperature and pressure conditions typical e.g. of the tropical tropopause (100 mbar, 220 K). As HOONO formation and loss are now parameterized (see above) in EMAC, we can explore its potential contribution to odd-nitrogen species in the atmosphere. The reaction between OH and NO_2 to form HOONO converts short lived HO*x* (HO*x* = OH + HO_2) and NO_X (NO_X = NO + NO_2) into a longer lived "reservoir" species, and in this sense is similar to the reaction between HO_2 and NO_2 to form HO_2NO_2

$$HO_2 + NO_2 + M \rightarrow HO_2NO_2 + M \qquad \text{(R9)}$$

which is also thermally unstable, dissociating to reform HO_2 and NO_2. Unlike HOONO, for which there are no atmospheric measurements, much effort has been made to measure concentrations of HO_2NO_2 in colder regions of the atmosphere and it is considered an important component of the NO_y budget at high altitudes (Nault et al., 2016). We therefore compared EMAC predictions of HOONO concentrations with those of HO_2NO_2. The results are displayed in Fig. 8, in which we plot the zonally averaged HOONO / HO_2NO_2 ratio. Immediately apparent from Fig. 8 is that, compared to HO_2NO_2, HOONO is a minor component of NO_y in the warm, lower atmosphere. This reflects the difference in the thermal decomposition rate constant of the two trace gases, that of HO_2NO_2 being ≈ 4×10^{-5} s^{-1} in e.g. the middle troposphere at 400 mbar and 250 K, whereas HOONO decomposes a factor 30 faster so that its lifetime is only ≈ 1000 s. In the UTLS region, the ratio increases further (HO_2NO_2 is a factor 50 more long-lived w.r.t. thermal decomposition at 100 mbar and 220 K) but the lifetimes of both gases under these conditions are sufficiently long that their concentrations are largely determined by their production rates and their losses due to photolysis and reaction with OH. The maximum ratio of HOONO to HO_2NO_2 is found at the tropical tropopause, where concentrations become comparable. As the modelled loss processes of HOONO and HO_2NO_2 (rate constants for photolysis and reaction with OH) are assumed to be identical, the occurrence of the maximum HOONO to HO_2NO_2 ratio at the tropical tropopause is related to the ratio of the (temperature dependent) rate coefficients responsible for their formation (at 220 K and 100 mbar this favours HOONO formation by a factor of ≈ 2) and the model OH / HO_2 ratio. Whilst this result

indicates that HOONO could be an important reservoir of NO_X under certain conditions, we must bear in mind that there is great uncertainty associated not only with the branching ratio to HOONO formation in R1b but also with its loss processes (reaction with OH, photolysis), which remain unexplored experimentally. OH reacts with HO_2NO_2 via H-abstraction from the H-OO group (IUPAC, 2019), and a similar mechanism is likely for HOONO. As the H-OO bond strength is likely to be greater in HOONO than in HO_2NO_2 (larger electron density around the peroxy bond) we may expect the rate coefficient to be lower for HOONO. A significantly lower rate coefficient for reaction with OH (or photolysis rate constant) could greatly increase the abundance of HOONO. If this were the case, airborne instruments that measure NO_x would likely also measure some fraction of HOONO following its rapid decomposition in warm inlet lines, as has been observed for HO_2NO_2 and $CH_3O_2NO_2$ (Nault et al., 2015; Silvern et al., 2018). Clearly, more experimental or theoretical data that better constrain the yield of HOONO and its atmospheric loss processes as well as atmospheric measurements are necessary in order to improve our understanding of the role of the reaction between OH and NO_2 throughout the atmosphere.

4 Conclusions

We have made very precise and accurate measurements for the overall rate coefficient, k_1, of the reaction between OH and NO_2, which is of critical importance in atmospheric chemistry. Our experiments demonstrate clearly that the presence of H_2O increases significantly the overall rate coefficient (k_1) of the reaction between OH and NO_2. H_2O is found to be a more efficient collisional quencher (by a factor of ≈ 6) of the initially formed HO-NO_2 association complex than N_2 and a factor ≈ 8 more efficient than O_2. A new parameterisation of the rate coefficient for the title reaction that considers the roles of N_2, O_2 and H_2O as third-body quenchers (also using data from our companion paper, Amedro et al. (2019)) has been incorporated into a global, chemistry transport model to assess its impact on e.g. the HNO_3 / NO_2 ratio as well as NO_X and OH levels. Compared to existing evaluations of the kinetic data, use of the new parameters will result in significant changes (5-10%) in the partitioning of NO_X and NO_Y, the direction of the bias depending on which evaluation is used as reference and on region of the atmosphere. This work highlights the continuing importance of obtaining accurate laboratory kinetic data for those reactions that are central to our understanding of atmospheric chemistry and which provide anchor-points in chemical transport models.

Though the result is associated with great uncertainty owing to missing kinetic parameters for HOONO, the global model predicts the presence of HOONO in concentrations similar to those of HO_2NO_2 at the tropical tropopause. The present dataset addresses only the overall rate coefficient, k_1. Detailed experimental studies of the formation of HOONO (e.g. its yield at various temperatures and in the presence of H_2O) and on the fate of HOONO (OH kinetics, photolysis) are required to better assess its role as NO_X and HO*x* reservoir in cold parts of the atmosphere.

Data availability. The rate coefficients measured during this experimental study are listed in Table 1.

Author contributions. The experiments were carried out by DA, AJCB, and MB. The data analysis and preparation of the paper were performed by DA, with assistance from JL and JNC. The global modelling was performed by AP.

Competing interests. The authors declare that they have no conflict of interest.

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Table 1. Measurements of k_1 in He bath-gas

T (K)	p (Torr)	M[a]	OH precursor	k_1[b]
277	48.6	1.68	H_2O_2	1.59 ± 0.14
292	25.1	0.83	H_2O_2 [c]	0.75 ± 0.07
	50.0	1.65	H_2O_2	1.37 ± 0.08
	75.1	2.47	H_2O_2	1.88 ± 0.12
	102.9	3.39	HNO_3	2.32 ± 0.15
	206.9	6.81	HNO_3 [d]	3.73 ± 0.25
	300.7	9.89	HNO_3	4.64 ± 0.29
	405.8	13.35	HNO_3	5.54 ± 0.37
	495.6	16.30	HNO_3	6.29 ± 0.40
	595.0	19.57	HNO_3	6.83 ± 0.42
	689.1	22.67	HNO_3	7.46 ± 0.46
332	28.1	0.82	H_2O_2	0.60 ± 0.06
	56.8	1.65	H_2O_2	0.99 ± 0.08
	85.4	2.48	H_2O_2	1.34 ± 0.10

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[a] Molecular density M(He) in units of 10^{18} molecule cm^{-3}. [b] Units of 10^{-12} cm^3 $molecule^{-1}$ s^{-1}. The errors are 2σ total uncertainty. [c]Concentration range of H_2O_2 ≈ 5-14 × 10^{13} molecule cm^{-3}. [d]Concentration range of HNO_3 ≈ 5-9 × 10^{13} molecule cm^{-3}.

Table 2. Measurements of k_1 in N_2-H_2O and He-H_2O bath-gas

T / K	p (Torr)	M[a]	$[H_2O]$[b]	x_{He} or x_{N2}	x_{H2O}	k_1[c]
N_2-H_2O bath gas						
292	50.2	1.65	0	1	0	2.58 ± 0.16
	50.2	1.66	0.86	0.950	0.050	3.07 ± 0.22
	50.0	1.65	1.62	0.905	0.095	3.45 ± 0.26
	50.0	1.65	2.28	0.866	0.134	3.83 ± 0.26
	50.2	1.66	2.84	0.834	0.166	3.95 ± 0.37
	49.2	1.63	3.27	0.805	0.195	4.10 ± 0.27
	50.0	1.65	4.06	0.754	0.246	4.47 ± 0.18 [d]
He-H_2O bath gas						
277	48.6	1.68	0	1	0	1.59 ± 0.11
	47.6	1.66	0.9	0.946	0.054	2.27 ± 0.15
	48.0	1.67	1.42	0.915	0.085	2.63 ± 0.17
	48.7	1.7	2	0.882	0.118	3.13 ± 0.24
291	50.0	1.65	0	1	0	1.37 ± 0.08
	50.6	1.68	0.64	0.962	0.038	1.99 ± 0.14
	51	1.69	1.30	0.923	0.077	2.39 ± 0.21
	50.7	1.68	2.25	0.863	0.137	2.88 ± 0.24
	49.5	1.64	3.06	0.818	0.182	3.43 ± 0.22
	50.8	1.68	3.12	0.810	0.190	3.44 ± 0.24
	49.7	1.65	3.60	0.783	0.217	3.54 ± 0.23
	50.2	1.66	3.94	0.764	0.236	3.72 ± 0.29
	50.5	1.67	4.68	0.721	0.279	4.08 ± 0.27
332	56.8	1.65	0	1	0	0.99 ± 0.06
	56.3	1.64	0.58	0.964	0.036	1.32 ± 0.08
	56	1.63	1.72	0.895	0.105	1.81 ± 0.16
	56.2	1.63	3.3	0.798	0.202	2.43 ± 0.18
	55.9	1.62	4.33	0.733	0.267	2.88 ± 0.22

Unless otherwise indicated, the measurements were performed using H_2O_2 as OH precursor. The concentration range of H_2O_2 was 5-18 × 10^{13} molecule cm^{-3} for experiments in He-H_2O bath gas and 9-14 × 10^{13} molecule cm^{-3} for experiments in N_2-H_2O bath gas. [a] Molecular density M(He-H_2O) or M(N_2-H_2O) in units of 10^{18} molecule cm^{-3}. [b]Units of 10^{17} molecule cm^{-3}. [c]Units of 10^{-12} cm^3 $molecule^{-1}$ s^{-1}. Errors are 2σ total uncertainty. [d] measurement performed using O_3-H_2O as OH precursor (with $[O_3]$ = 2 × 10^{13} molecule cm^{-3}).

Table 3. Parameters for calculating k_1 using Eqn. (15) and (16)

Bath-gas	k_0[a]	T-dependence of k_0 (m, q or o)	k_∞[b]	F_c
N_2	2.6×10^{-30}	3.6 (m)	6.3×10^{-11}	0.39
O_2	2.0×10^{-30}	3.6 (q)		
H_2O	15.9×10^{-30}	3.4 (o)		

620

[a] Units of cm^6 $molecule^{-2}$ s^{-1}. [b] Units of cm^3 $molecule^{-1}$ s^{-1}. Note that k_∞ is independent of temperature (n = 0).

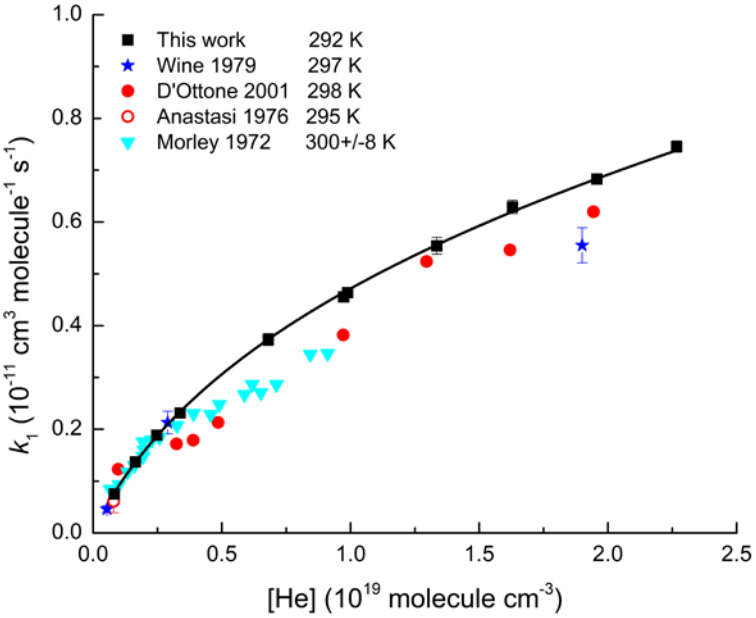

Figure 1. Values of k_1 from this study (black squares) as a function of He concentration at 292 K. Errors are 2σ statistical only. The solid line is a fit to our data using Eqn. (4) with $k_0 = 1.4 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-1} s^{-1}, $F_c = 0.32$, $m = 3.1$ and $n = 0$. Previous datasets at room temperature (Wine et al. (1979), D'Ottone et al. (2001), Anastasi and Smith (1976) and Morley and Smith (1972)) are displayed for comparison.

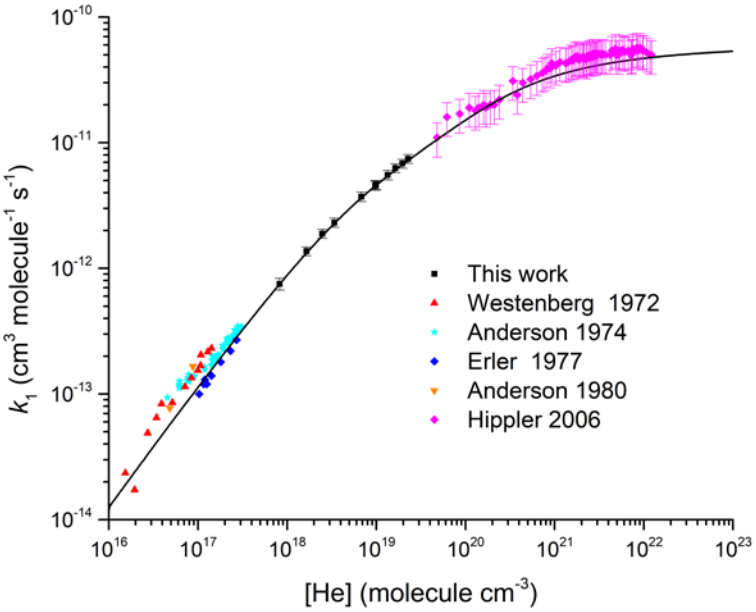

Figure 2. Comparison between the present dataset, the high pressure measurements by Hippler et al. (2006) and the low pressure measurements by Anderson et al. (1974), Westenberg and Dehaas (1972), Anderson (1980) and Erler et al. (1977). All measurements were made at room-temperature. The black line is our parameterisation with $k_0 = 1.4 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-1} s^{-1}, $m = 3.1$, $n = 0$ and $F_c = 0.32$.

625

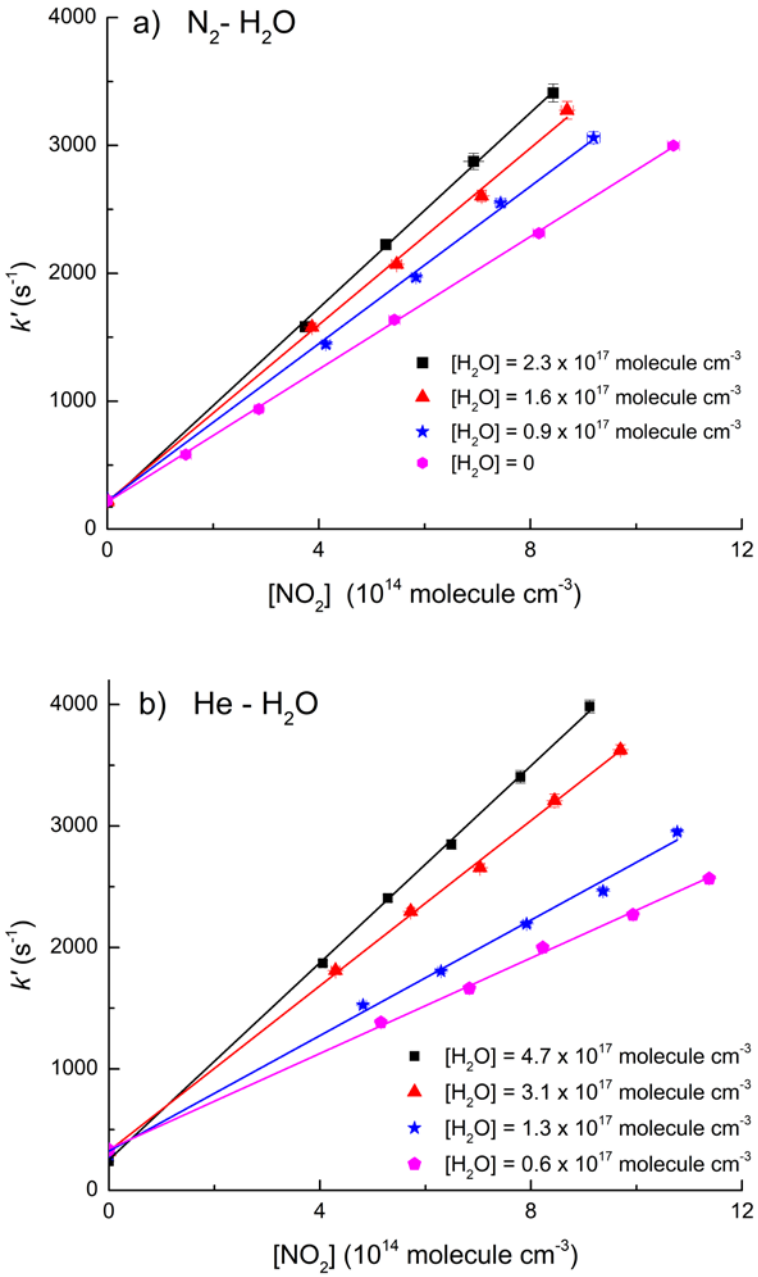

Figure 3 a) Data obtained in N_2-H_2O bath-gas (50 Torr, 292 K). **b)** Data obtained in He-H_2O bath-gas (50 Torr, 291 K). Both panels display first-order, OH decay constants in various concentrations of NO_2 and different mole fractions of H_2O. The solid lines represent least squares linear fits to Eqn. (2).

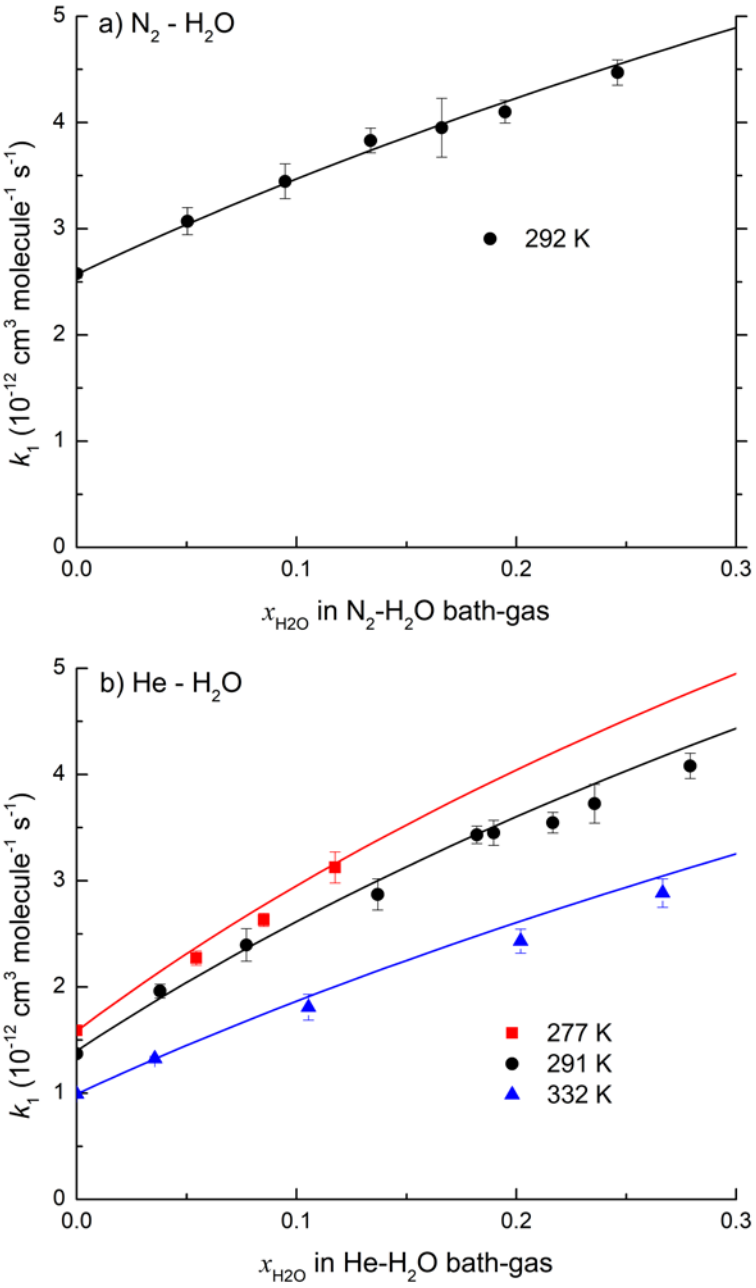

Figure 4. a) k_1 as a function of x_{H2O} at 50 Torr N_2-H_2O and 292 K. The line represents a least squares, multivariate fit (Eqn. 7 and 8) with $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-2} s^{-1}, $k_0^{N2} = 2.6 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $F_c = 0.39$, $m = 3.6$, $k_0^{H2O} = 15.9 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $o = 3.4$. **b)** k_1 as a function of x_{H2O} in He-H_2O mixtures at 277, 291 and 332 K. The solid lines represent a least squares, multivariate fit (Eqn. 7 and 9 to 12) where $k_\infty = 6.3 \times 10^{-11}$ cm^3 molecule^{-2} s^{-1}, $k_0^{He} = 1.4 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $F_c^{He} = 0.32$, $m = 3.1$, $k_0^{H2O} = 15.9 \times 10^{-30}$ cm^6 molecule^{-2} s^{-1}, $F_c^{H2O} = 0.39$ and $o = 3.4$.

630

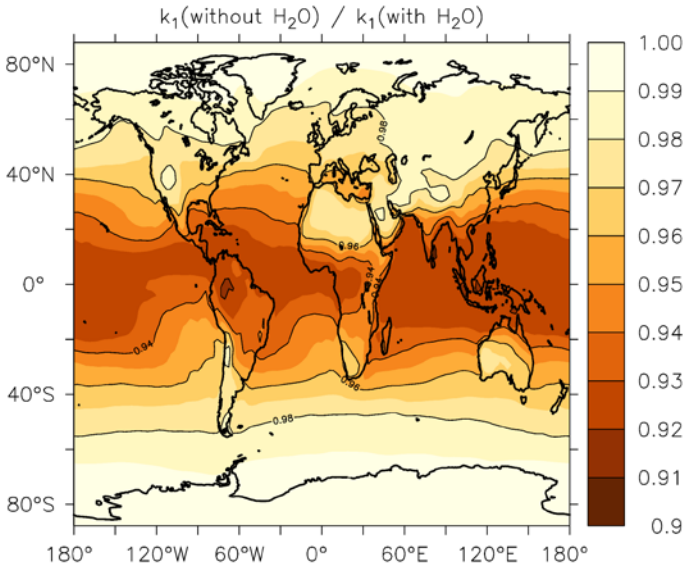

Figure 5. Annual average effect of H_2O on k_1 expressed as the fractional change in the rate coefficient at the Earth's surface when setting the mole fraction of water vapour to zero in Eqn. 15.

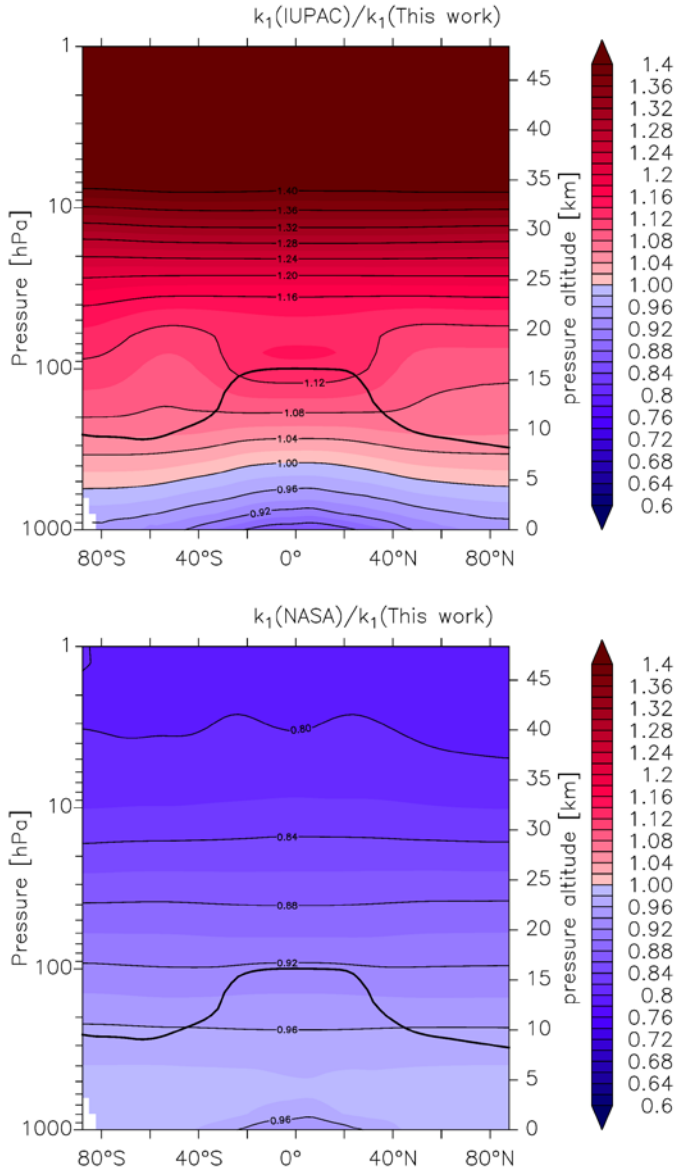

Fig. 6 Global values of $\frac{k_1^{\mathrm{IUPAC}}}{k_1^{\mathrm{this\ work}}}$ (upper panel) and $\frac{k_1^{\mathrm{NASA}}}{k_1^{\mathrm{this\ work}}}$ lower panel). k_1 is the overall rate coefficient (both channels) for Reaction R1 calculated using the parameters from this work ($k_1^{\mathrm{this\ work}}$) and those presently recommended by the IUPAC (k_1^{IUPAC}) and NASA (k_1^{NASA}) data evaluation panels. The black line represents the model tropopause.

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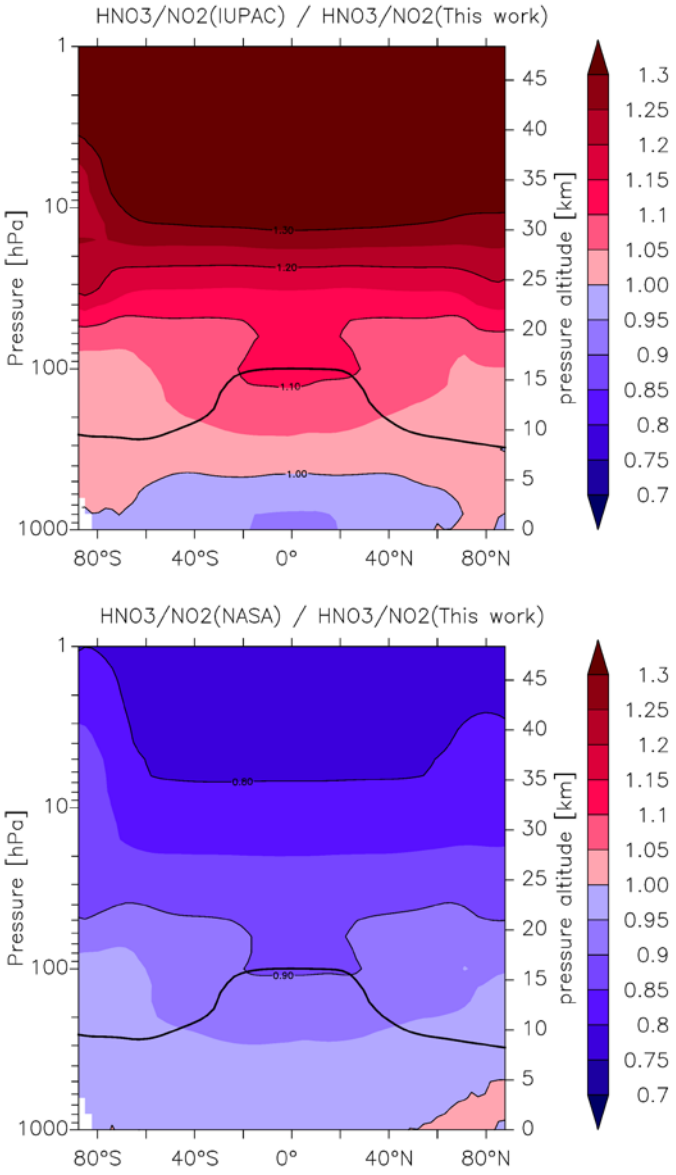

Figure 7. Effect of different parameterisations of k_1 on the global (zonal and yearly averaged) HNO_3 to NO_2 ratio. The upper panel plots $\frac{HNO3}{NO2}$ (IUPAC)/ $\frac{HNO3}{NO2}$ (this work), the lower panel plots $\frac{HNO3}{NO2}$ (NASA)/ $\frac{HNO3}{NO2}$ (this work). The black line represents the model tropopause.

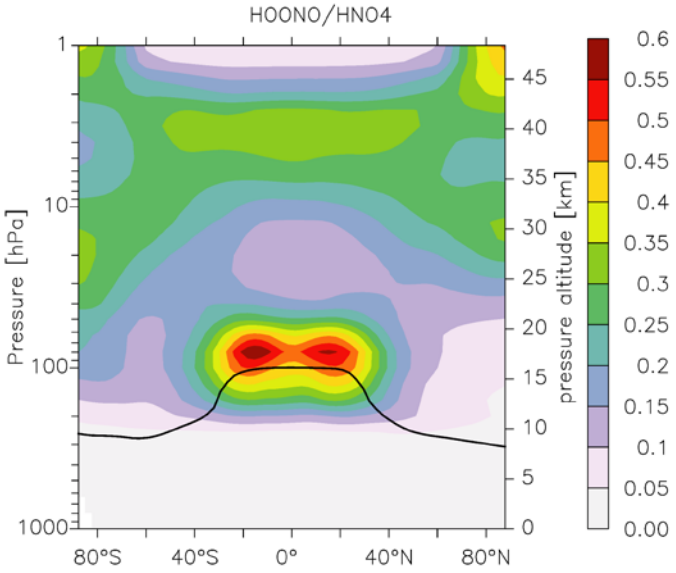

Figure 8. Model (EMAC) ratio of HOONO (formed in the reaction of NO_2 with OH) to HO_2NO_2 (formed in the reaction of NO_2 with HO_2) calculated using the present parameterisation of k_1 and equating the (unknown) rate coefficients for loss of HOONO via reaction with OH or photolysis to those of HO_2NO_2. The black line represents the model tropopause.

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