

Interactive comment on “Technical note: Determination of binary gas phase diffusion coefficients of unstable and adsorbing atmospheric trace gases at low temperature – Arrested Flow and Twin Tube method” by Stefan Langenberg et al.

Stefan Langenberg et al.

langenberg@uni-bonn.de

Received and published: 22 January 2020

We thank the referee for reviewing and commenting our discussion paper. The remarks of the reviewer are marked like *this*. All symbols and equations used and cited herein refer to the discussion paper unless otherwise indicated. To the comments and questions we answer as follows:

C1

In the experimental part, the authors shortly mention the limiting cases for each of the two methods. How many wall collisions do the diffusing molecules typically undergo?

The collision rate per unit area is given by

$$Z_w = \frac{1}{4} c \bar{v}$$

where

$$\bar{v} = \sqrt{\frac{8kT}{\pi m}} = \sqrt{\frac{8RT}{\pi M}}$$

is the mean molecular velocity of the species under investigation. If the trace gas contact time τ is given, the (dimensionless) number of hits of an individual species molecule on the cylindrical surface can be estimated

$$N_{\text{hit}} = \frac{\tau \bar{v}}{2r}.$$

For the AF-method τ corresponds to the arrest time t_a . For the TT-method τ is the mean travel time of a molecule along the length l of the capillary. This yields

$$\tau = \frac{l^2}{2D}.$$

For ozone at 273.15 K $\tau = 14.4$ s is found for the TT-method. Thus, the number of hits is 9×10^6 for the AF-method at a arrest time of 100 s and 6×10^7 for the TT-method.

Since either reversible adsorption or chemical reaction are affecting the transport kinetics, the authors could elaborate the limiting first order loss rate coefficient and the limiting residence time on the surface to lead to a noticeable impact on the analysis of the detector signals for each method.

The AF-method is not affected by first order loss. However, most heterogeneous loss processes are not strictly first order. The TT-method *is* affected by lost processes. A

C2

tolerable upper limit for the uptake/reaction coefficient γ for enabling the TT-method is derived as follows: let's assume that the ratio between the reactive flow J_R to the surface and the diffusion flow through the capillary should not exceed 0.01. Thus, $J_R < 0.01J_D$. The reactive flow the surface is given by

$$J_R = \frac{\bar{v}\gamma}{4l} \int_0^l \tilde{c}(z) dz,$$

considering that the concentration \tilde{c} is a linear function of the coordinate z along the capillary tube. We obtain

$$\tilde{c}(z) = c_0 + \frac{z}{l}(c - c_0).$$

This yields

$$J_R = \frac{1}{8} \bar{v}\gamma(c_0 + c).$$

Thus, it follows with Eq. (16)

$$\gamma < 0.08 \frac{D}{\bar{v}l} \left(\frac{c_0 - c}{c_0 + c} \right) \approx 0.08 \frac{D}{\bar{v}l}.$$

For ozone at 273.15 K with $\bar{v} = 347.1 \text{ m s}^{-1}$ $\gamma \ll 1.7 \times 10^{-7}$ is required for the successful application of the TT-method. Ozone destruction on similar quartz surfaces was already studied by Langenberg and Schurath (1999): $\gamma = 4.4 \times 10^{-7}$ was found for ozone at room temperature. Therefore, the TT-method is not suitable to determine the diffusion coefficient of ozone. In addition, further loss of ozone in other parts of the apparatus needs to be considered too.

Could the method in turn be used to measure the surface residence time of sticky but non-reactive molecules through their effective diffusivity?

The TT-method is not affected by non-reactive adsorption. However, the AF-method may be affected. The most simplistic model for adsorption is Henry's adsorption

C3

isotherm: the surface concentration q as function of the trace gas partial pressure p is given by (Langenberg and Schurath, 2018)

$$q = K_H p.$$

The capacity ratio is defined as

$$k' = \frac{2RT}{r} K_H.$$

In case of adsorption it follows from the mass balance in the diffusion capillary

$$\left(\frac{\partial c}{\partial t} \right)_z = \frac{D}{1 + k'} \left(\frac{\partial^2 c}{\partial z^2} \right)_t.$$

Therefore, the diffusion coefficient measured by the AF-method would be smaller than the actual diffusion coefficient. Thus, if the real diffusion coefficient D is known, measurement of the effective diffusion coefficients can be used to determine k' and K_H . However, it is much easier to determine k' using standard gas chromatography by measuring arrest times (first central moment) instead of measurement of peak broadening (second central moment).

Line 190: no need to decide whether ozone is adsorbing or non-adsorbing. Each molecule may adsorb. I suggest to simply mention chemical decay. Whether it undergoes reversible adsorption in addition seems not obvious (cf also previous comment) from the data but cannot be excluded.

We also tried to investigate the diffusion of NO_2 using the AF-method. However, adsorption gave rise to strong peak tailing which invalidated the AF-method.

References

Langenberg, S. and Schurath, U.: Ozone Destruction on Ice, *Geophys. Res. Lett.*, 26, 1695–1698, <https://doi.org/10.1029/1999GL900325>, 1999.

Langenberg, S. and Schurath, U.: Gas chromatography using ice-coated fused silica columns: study of adsorption of sulfur dioxide on water ice, *Atmos. Chem. Phys.*, 18, 7527–7537, <https://doi.org/10.5194/acp-18-7527-2018>, 2018.

Interactive comment on *Atmos. Chem. Phys. Discuss.*, <https://doi.org/10.5194/acp-2019-1050>, 2019.

Interactive comment on “Technical note: Determination of binary gas phase diffusion coefficients of unstable and adsorbing atmospheric trace gases at low temperature – Arrested Flow and Twin Tube method” by Stefan Langenberg et al.

Stefan Langenberg et al.

langenberg@uni-bonn.de

Received and published: 18 February 2020

We thank the referee for reviewing and commenting our discussion paper. The remarks of the reviewer are marked like *this*. All symbols, equations and references used and cited herein refer to the discussion paper unless otherwise indicated. To the comments and questions we answer as follows:

C1

Something in the big picture is missing: the application to the laboratory kinetics measurements whose results might depend on the accuracy of the diffusion coefficients (the stated reason for this detailed work, lines 60-62.) How will the results of these new measurement capabilities affect previously measured uptake coefficients? It seems that the ClONO₂ and N₂O₅ diffusion coefficients might have the most impact in this regard. A recommended set of L-J parameters for these two species would be most interesting.

Most uptake experiments with ClONO₂ and N₂O₅ into liquid surfaces were performed by the droplet train technique. To determine the real uptake coefficient γ_0 , the measured uptake coefficient γ_{obs} had to be corrected for diffusion to the droplet surface. This can be performed by a simplistic resistance model of Hu et al. (1995)

$$\frac{1}{\gamma} = \frac{1}{\gamma_{\text{obs}}} - \frac{\bar{v}r}{4D_g} + \frac{1}{2}$$

where \bar{v} is the average trace gas thermal velocity and r the particle radius. The necessary diffusion coefficients were mostly taken from Hanson and Ravishankara (1991), who estimated them using the Lennard-Jones model and parameters of Patrick and Golden (1983) listed in our Table 1. In general, the investigators did not perform a sensitivity analysis of the dependency of the diffusion coefficient upon their reported γ . Therefore, it is difficult to assess, how a change of the diffusion coefficient would change the value of γ obtained.

We try to estimate this interdependency using the work of George et. al. (1994), who studied the uptake of N₂O₅ into water droplets. The dependency of the relative error of γ on the relative error of the diffusion coefficient D_g is given by

$$\frac{\Delta\gamma}{\gamma} = \frac{D_g}{\gamma} \left(\frac{\partial\gamma}{\partial D_g} \right) \frac{\Delta D_g}{D_g}.$$

C2

Using the resistance model for a spherical particle, this yields

$$\frac{\Delta\gamma}{\gamma} = \frac{\gamma\bar{v}r}{4D_g} \frac{\Delta D_g}{D_g}.$$

George et. al. (1994) calculated the diffusion coefficient of N_2O_5 in N_2 using the method of Fuller et al. (1966). Back calculated to 273.15 K, they obtained $D_0 = 0.112 \text{ cm}^2\text{s}^{-1}$. This value is about 24% higher than the diffusion coefficient $0.09 \text{ cm}^2\text{s}^{-1}$ estimated by the LJ-model, which is consistent with our data within the error limits. At 273 K and 25.2 Torr, they obtained $\gamma = 0.020 \pm 0.002$. Assuming a particle radius of $60 \mu\text{m}$, a diffusion coefficient of $D = 3.38 \text{ cm}^2\text{s}^{-1}$ at 25.2 Torr and $\bar{v} = 231.4 \text{ m s}^{-1}$, the relative error is $\Delta\gamma/\gamma = 0.21\Delta D_g/D_g = 5\%$.

To estimate γ for the uptake of ClONO_2 on ice and subsequent reaction, the diffusion coefficient might be of greater importance. Hanson and Ravishankara (1992) reported for their flow tube reactor study, gas phase diffusion limits transport to the ice surface. They were only able to report an upper limit of $\gamma > 0.3$ since the diffusion coefficient of ClONO_2 in He was not accurately known.

We conclude that the diffusion coefficients of ClONO_2 and N_2O_5 calculated by the LJ-model using σ and ϵ from Patrick and Golden (1983) are a good choice. Therefore, it does not make sense to backward calculate σ and ϵ from our experimental data.

The concerns about the measurements center around these two molecules and the temperature dependencies of the D's. Both N_2O_5 's and ClONO_2 's measured T-dependencies differ significantly from that expected for L-J interactions.

We use the `anova()` test of R (Phillips, 2018) to check if the temperature coefficient b in our Table 3 which we determined for ClONO_2 and N_2O_5 significantly deviates from the temperature coefficient of the LJ-model. To do this, first a fit is performed using Eq. (5) and Eq. (24) with D_0 and b fit parameters. This yields the values given in Table 3. Then the fit is repeated setting b to the value of the Lennard-Jones model. Now the

C3

two fits (statistical models) are compared using the `anova()` test. As you can see, the T-dependencies does not differ significantly from b predicted by Eq. (7):

Species	Carrier	b (LJ)	D_0 [cm^2s^{-1}]	a [s]	P	Deviation
ClONO_2	He	1.72	0.307 ± 0.003	–	0.37	not significant
ClONO_2	N_2	1.88	0.086 ± 0.001	–	0.08	weakly significant
N_2O_5	He	1.73	0.31 ± 0.01	-12 ± 3	0.16	not significant
N_2O_5	N_2	1.91	0.085 ± 0.002	-29 ± 4	0.35	not significant

While losses were addressed, it seems these anomalous T-dependencies suggest there is more to the story. The indirect detection method for these two species is worth some consideration. Hard to come up with reasons why these two molecules interacting with He and N_2 should not be describable by L-J potentials.

We don't think that the titration reaction with NO can affect the temperature dependency because the titration reaction is always performed outside the thermostatted cold box in a heating zone.

Regarding the LJ-model, two things have to be considered:

1. ClONO_2 and N_2O_5 are both polar components with a permanent dipole moment. Therefore, the interactions with He and N_2 are not only van der Waals interactions but also dipole – induced dipole interactions. The latter are not considered in the simple LJ-model.
2. For unstable compounds like ClONO_2 and N_2O_5 ϵ cannot be obtained from viscosity data. ϵ can only be estimated from the boiling point using Eq. (4). But for polar component this only can be a rough estimation since the boiling point not only depends on the van der Waals interactions but also on dipole – dipole interactions.

C4

References

George, C., Ponche, J. L., Mirabel, P., Behnke, W., Scheer, V., and Zetzsch, C.: Study of the Uptake of N_2O_5 by Water and NaCl Solutions, *J. Phys. Chem.*, 98, 8780–8784, <https://doi.org/10.1021/j100086a031>, 1994.

Hanson, D., Ravishankara, A.: The Reaction Probabilities of ClONO_2 and N_2O_5 on Polar Stratospheric Cloud Materials, *J. Geophys. Res.*, 96, 5081–509, <https://doi.org/10.1029/90JD02613>, 1991

Hanson, D., Ravishankara, A.: Investigation of the reactive and nonreactive processes involving nitryl hypochlorite and hydrogen chloride on water and nitric acid doped ice, *J. Phys. Chem.*, 96, 2682–2691, <https://doi.org/10.1021/j100185a052>, 1992

Hu, J.H., Shi, Q., Davidovits, P., Worsnop, D.R., Zahniser, M.S. and Kolb C.E.: Reactive Uptake of $\text{Cl}_2(\text{g})$ and $\text{Br}_2(\text{g})$ by Aqueous Surfaces as a Function of Br^- and Cl^- Concentration: The Effect of Chemical Reaction at the Interface, *J. Phys. Chem.*, 99, 8768, <https://doi.org/10.1021/j100021a050>, 1995.

Phillips, N.D.: YaRrr! The Pirate's Guide to R, <https://bookdown.org/ndphillips/YaRrr/comparing-regression-models-with-anova.html> , 2018, (last visited February 18, 2020)

Interactive comment on Atmos. Chem. Phys. Discuss., <https://doi.org/10.5194/acp-2019-1050>, 2019.

Technical note: Determination of binary gas phase diffusion coefficients of unstable and adsorbing atmospheric trace gases at low temperature – Arrested Flow and Twin Tube method

Stefan Langenberg^{1,a}, Torsten Carstens^{1,2}, Dirk Hupperich¹, Silke Schweighoefer^{1,b}, and Ulrich Schurath³

¹Institut für Physikalische und Theoretische Chemie, University of Bonn, Bonn, Germany

²Karlsruher Institut für Technologie, Karlsruhe, Germany

³Institut für Umweltphysik, University of Heidelberg, Heidelberg, Germany

^anow at: Klinik und Poliklinik für Hals-Nasen-Ohrenheilkunde/Chirurgie, University of Bonn, Bonn, Germany

^bnow at: UP GmbH, Ibbenbüren, Germany

Correspondence: Stefan Langenberg (langenberg@uni-bonn.de)

Abstract. Gas phase diffusion is the first step for all heterogeneous reactions under atmospheric conditions. Knowledge of binary diffusion coefficients is important for the interpretation of laboratory studies regarding heterogeneous trace gas uptake and reactions. Only for stable, nonreactive and non polar gases well-established models for the estimation of diffusion coefficients from viscosity data do exist. Therefore, we have used two complementary methods for the measurement of binary diffusion coefficients in the temperature range of 200 K to 300 K: the arrested flow method is best suited for unstable gases and the twin tube method is best suited for stable but adsorbing trace gases. Both methods were validated by measurement of diffusion coefficients of methane and ethane in helium and air and nitric oxide in helium. Using the arrested flow method the diffusion coefficients of ozone in air, dinitrogen pentoxide and chlorine nitrate in helium and nitrogen were measured. The twin tube method was used for the measurement of the diffusion coefficient of nitrogen dioxide and dinitrogen tetroxide in helium and nitrogen.

Copyright statement. This work is distributed under the Creative Commons Attribution 4.0 License.

1 Introduction

The critical role of heterogeneous reactions in atmospheric chemistry is widely accepted. The diffusion of gas molecules towards the surface is the first step in a heterogeneous reaction, and it can influence and sometimes even control the overall rate of the uptake of a trace gas onto the surface (Kolb et al., 2010; Tang et al., 2014a). Diffusion also plays a role in atmosphere – biosphere interactions: the incorporation of trace gases like ozone and nitrogen dioxide into leaves and isoprene out through stomata is diffusion controlled (Laisk et al., 1989; Eller and Sparks, 2006; Fall and Monson, 1992).

Table 1. Lennard-Jones parameters of the species investigated in this study

Species	Formula	M [g mol ⁻¹]	σ [Å]	ϵ/k [K]	Method	Source
Helium	He	4.00	2.551	10.2	v	Poling et al. (2004)
Nitrogen	N ₂	28.01	3.798	71.4	v	Poling et al. (2004)
Air		28.81	3.711	78.6	v	Poling et al. (2004)
Methane	CH ₄	16.04	3.758	148.6	v	Poling et al. (2004)
Ethene	C ₂ H ₄	28.05	4.163	224.7	v	Poling et al. (2004)
Nitric oxide	NO	30.01	3.492	116.7	v	Poling et al. (2004)
Nitrogen dioxide	NO ₂	46.01	3.765	210.0	v	Brokaw and Svehla (1966)
Ozone	O ₃	48.00	3.875	208.4	b	Massman (1998)
Dinitrogen tetroxide	N ₂ O ₄	92.01	4.621	347.0	v	Brokaw and Svehla (1966)
Chlorine nitrate	ClONO ₂	97.46	4.470	364.7	b	Patrick and Golden (1983)
Dinitrogen pentoxide	N ₂ O ₅	108.01	4.570	450.0	b	Patrick and Golden (1983)

v = obtained from viscosity data, b = obtained from T_b and V_b using Eqns. (4).

Marrero and Mason (1972); Massman (1998); Tang et al. (2014a, 2015) and Gu et al. (2018) compiled and evaluated the available experimental data of diffusion coefficients of atmospheric trace gases. However, the existing compilations focus on stable gases, experimental diffusion coefficients of ozone, nitrogen dioxide, chlorine nitrate and dinitrogen pentoxide are still missing. They cannot be predicted with the required accuracy because detailed kinetic theory requires intermolecular potentials which are not generally available for atmospherically relevant compounds.

Chapman and Enskog derived the following equation from kinetic theory of gases for the molecular binary diffusion coefficient

$$D = \frac{3}{16} \sqrt{\frac{2\pi kT(m_A + m_B)}{m_A m_B}} \left(\frac{kT}{\pi \sigma_{AB}^2 \Omega_D p} \right) \quad (1)$$

where m is the mass of the molecules, k is the Boltzmann constant, p the pressure and T is the absolute temperature. σ_{AB} is the characteristic length of the intermolecular force law, Ω_D is the dimensionless collision integral of diffusion. It depends on the temperature and the characteristic energy ϵ_{AB} of the Lennard-Jones potential describing the intermolecular force (Poling et al., 2004; Marrero and Mason, 1972). Ω_D as function of temperature is expressed by the fit function

$$\Omega_D = \frac{A}{\Theta^B} + \frac{C}{\exp(D\Theta)} + \frac{E}{\exp(F\Theta)} + \frac{G}{\exp(H\Theta)} \quad (2)$$

where $\Theta = kT/\epsilon_{AB}$, $A = 1.06036$, $B = 0.15610$, $C = 0.19300$, $D = 0.47635$, $E = 1.03587$, $F = 1.52996$, $G = 1.76474$, $H = 3.89411$ (Neufeld et al., 1972; Poling et al., 2004). The equations

$$\epsilon_{AB} = \sqrt{\epsilon_A \epsilon_B}, \quad \sigma_{AB} = \frac{\sigma_A + \sigma_B}{2} \quad (3)$$

are usually employed to relate the interaction parameters of the Lennard-Jones potential between components A and B to the interaction potential parameters of the individual components. A tabulation of the potential parameters of the species considered in this work is given in Table 1. The Lennard-Jones parameters σ and ϵ are generally not available for unstable atmospheric trace gases. Patrick and Golden (1983) estimated them by the equations

$$\sigma = 1.18 V_b^{1/3}, \quad \epsilon/k = 1.21 T_b \quad (4)$$

from T_b the normal boiling point temperature and V_b the molar volume at boiling point. In cases where V_b cannot be determined experimentally, it is obtained from tables of atomic volumes using the LeBas method. Patrick and Golden (1983) assumed the systematic errors of σ and ϵ obtained by this method to be $\leq 20\%$.

The diffusion coefficient as function of pressure in a narrow temperature range close to the reference temperature T_0 is usually expressed as

$$D = D_0 \left(\frac{p_0}{p} \right) \left(\frac{T}{T_0} \right)^b \quad (5)$$

where $T_0 = 273.15$ K are the standard temperature and $p_0 = 101325$ Pa the standard pressure (STP). Close to the reference temperature T_0 , the temperature coefficient b can be calculated as follows (Poling et al., 2004)

$$b = \left(\frac{\partial \ln D}{\partial \ln T} \right) = \frac{3}{2} - \left(\frac{\partial \ln \Omega_D}{\partial \ln T} \right) = \frac{3}{2} - \frac{T_0}{\Omega_D} \left(\frac{\partial \Omega_D}{\partial T} \right), \quad (6)$$

$$b = \frac{3}{2} - \frac{\Theta}{\Omega_D} \left(\frac{\partial \Omega_D}{\partial \Theta} \right) \quad (7)$$

From Eq. (2) it is obtained by derivation

$$\left(\frac{\partial \Omega_D}{\partial \Theta} \right) = -\frac{AB}{\Theta^{B+1}} - \frac{CD}{\exp(D\Theta)} - \frac{EF}{\exp(F\Theta)} - \frac{GH}{\exp(H\Theta)} \quad (8)$$

Fuller et al. (1966) developed a simple correlation equation for the estimation of gas phase diffusion coefficients using additive atomic volumes V_A and V_B for each species. With the molar masses M_A and M_B ($[M] = \text{g mol}^{-1}$) of each species and $[p] = \text{bar}$, the diffusion coefficient ($[D] = \text{cm}^2 \text{s}^{-1}$) is given by

$$M_{AB} = \frac{2}{1/M_A + 1/M_B} \quad (9)$$

$$D = 0.00143 \frac{T^{1.75}}{\sqrt{M_{AB}}(V_A^{1/3} + V_B^{1/3})^2 p} \quad (10)$$

Tabulations of atomic volume increments are summarized by Poling et al. (2004) and Tang et al. (2014a).

In the atmosphere, for typical submicron-sized aerosol particles, gas-phase diffusion does not usually limit uptake. Therefore, for modelling atmospheric processes, it is sufficient to use diffusion coefficients obtained using the Fuller method. However, in many laboratory experiments for the measurement of mass accommodation coefficients, conditions are such that gas-phase diffusion limitations need to be taken into account (Kirchner et al., 1990; Müller and Heal, 2002; Davidovits et al., 2006).

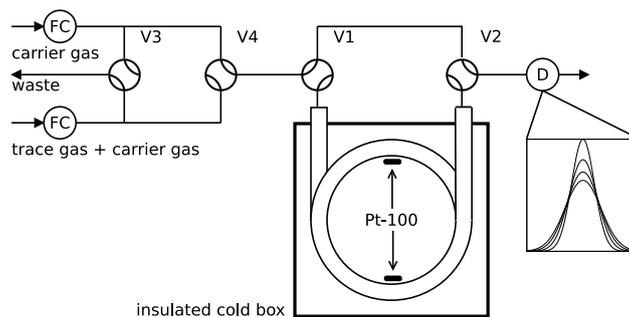


Figure 1. Arrested flow method: a pulse of trace gas is introduced into the column by simultaneously switching valves V3 and V4 for a short time. After the peak has reached the middle of the column, the carrier gas is bypassed by switching V1 and V2 for arrest times t_a between 0 and 200 s. After each arrest time t_a the corresponding peak shape is recorded by the detector D. The gas flow is controlled by flow controller FC.

2 Methods

2.1 Arrested flow method

- 65 The arrested flow (AF) method was first described by Knox and McLaren (1964) and McCoy and Moffat (1986): the diffusion coefficient of a given trace gas is derived from the broadening of width ζ_t of trace gas plugs arrested for different times in a long void gas chromatography glass column (length $l = 2.8$ m, radius $r = 0.189$ cm). A plug is generated by injecting a small amount of dilute trace gas into a steady flow of carrier gas by means of computer controlled solenoid valves. The flow is arrested when the plug has travelled halfway down the tube, see Fig. 1. In the absence of turbulence, the initial plug profile
- 70 spreads out along the tube by molecular diffusion only. Until the flow is arrested, the box profile of the trace gas has reshaped to Gaussian by Taylor diffusion (Taylor, 1953, 1954), if the condition

$$l \gg \frac{\dot{V}}{\pi D} \quad (11)$$

is fulfilled, where \dot{V} is the carrier gas flow rate. After a given arrest time t_a , the trace gas is eluted with ≈ 20 sccm (1 sccm = 1 ml min⁻¹ at 273.15 K and 1013 hPa) and the concentration profile is measured with a suitable gas chromatography detector.

- 75 This procedure is repeated for different arrest times t_a . The experimental peak profiles are fitted to Gaussians to determine the peak variance ζ_t^2 . According to theory based on Fick's second law of diffusion

$$\left(\frac{\partial c}{\partial t}\right)_z = D \left(\frac{\partial^2 c}{\partial z^2}\right)_t \quad (12)$$

a plot of ζ_t^2 versus arrest time t_a should be linear. The slope of the plot of ζ_z vs. t_a is given by

$$\frac{\Delta \zeta_z^2}{\Delta t_a} = 2D \quad (13)$$

80 Since the variance is measured in units of time, it has to be converted to units of length using the gas flow speed v in the column

$$\Delta\zeta_z^2 = v^2 \Delta\zeta_t^2. \quad (14)$$

From the carrier gas mass flow \dot{n} , temperature T and pressure p in the column which approximately equals atmospheric pressure, the flow speed can be determined by

85
$$v = \frac{\dot{n}RT}{\pi r^2 p} \quad (15)$$

The column is embedded in an aluminum block which is cooled by a recirculating cryostat (Lauda RLS6). The aluminum block is mounted in a plastic box insulated by Styrodur. The column temperature homogeneity is monitored with 2 Pt-100 sensors connected to the upper and lower parts of the column coil. The solenoid valves are connected by 1/16" Teflon tubes and controlled by computer using the software Asyst 3.1 (Keithley). At each temperature, 12 to 20 peaks are recorded at different
90 arrest times.

The systematic error of the determined diffusion coefficients using this method primarily depends on the systematic error of measuring the inner diameter of the column and the systematic error of the mass flow rate. A Teflon tube pushed through the column was used to determine the length of the column. The void volume of the column was determined by filling the column with water and measuring the weight of the water. From volume and length, the cross-sectional area and radius are calculated,
95 yielding a mean radius with a systematic error of 0.5%. After the experiments, the column was cut into small fragments. The inner diameter of these fragments was measured using a caliper gauge. We found that the inner diameter synchronously changes with column winding with a variability of 1%. When using Eq. (14) to transform $\Delta\zeta_t$ to $\Delta\zeta_z$, not the mean velocity but the actual velocity v and radius r at the location where the peak is arrested are relevant. Therefore, the actual systematic error of the radius is about 1%. The mass flow controllers were calibrated using a soap bubble flow meter. Thus, the systematic
100 error of the mass flow rate is about 1.5%. This sums up to a total theoretical systematic error of the AF method to about 7%. The random error of the method is about $> 0.4\%$, twice the repeatability $> 0.2\%$ of the flow rate.

2.2 Twin tube method

The twin tube (TT) method is a steady state technique for diffusion coefficient measurements over a wide temperature range using a diffusion bridge (Marrero and Mason, 1972). It is insensitive to wall adsorption effects which may invalidate AF
105 measurements at low temperature. Our apparatus consists of two parallel horizontal flow tubes (length 2 m, inner diameter 10 mm) which are connected by a bunch of $n = 220$ carefully thermostatted fused silica capillaries of radius $r = (39.2 \pm 0.4) \mu\text{m}$ and length $l = (20.8 \pm 0.3) \text{mm}$, see Fig. 2. The capillaries are embedded in a block made of brass. The cooling liquid of a cryostat (Lauda RLS6) circulates through the brass block, thereby covering the range from ambient temperature down to 198 K. Close to the diffusion bridge, the temperature in the block is measured with two Pt-100 sensors. The capillaries are pasted into
110 two parallel slits in a short section of the parallel flow tubes that is made of stainless steel. Up and downstream of the brass block, the flow tubes consist of glass. The entire apparatus consisting of the flow tubes and the diffusion bridge are housed

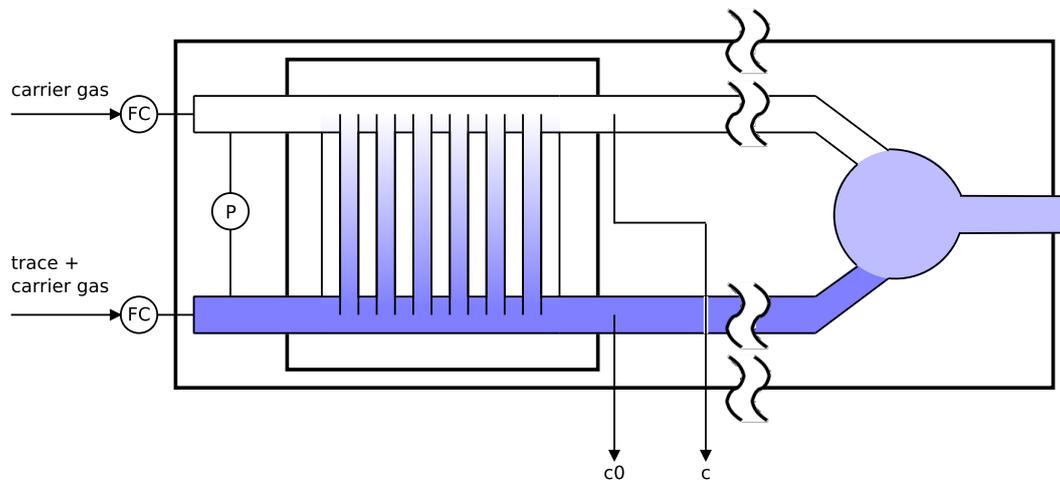


Figure 2. Schematic diagram of the twin tube experiment (false to scale): a diffusion bridge is connecting two flow tubes. Downstream the diffusion bridge some gas is continuously sampled for analysis to determine the trace gas concentrations c_0 and c .

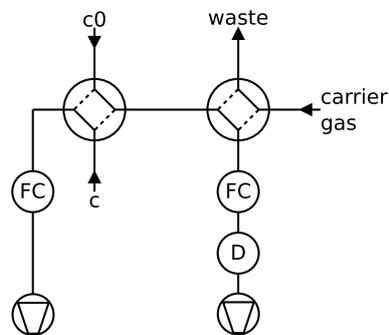


Figure 3. Twin tube method, continuous mode: a partial current is sucked through the detector D by a mass flow controller FC. A second 4-port valve enables to switch the detector to pure carrier gas to record the baseline.

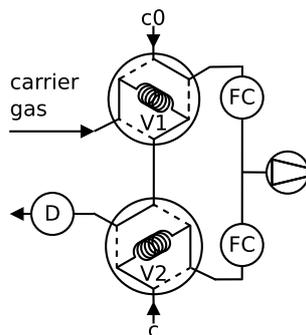


Figure 4. Twin tube method, peak mode: the carrier gas is admitted first through the sample loop V1 or V2 of a six-port valve. Then the content of the sample loop is pushed into the detector as peak after switching the six-port valve.

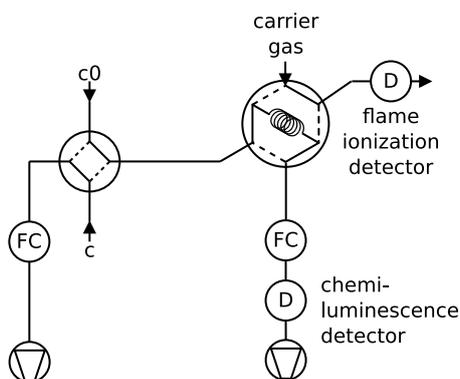


Figure 5. Twin tube method with internal standard as used for NO_2 in N_2 : the species under investigation is monitored in continuous mode, the internal standard is sampled by a 6-port valve and detected by a flame ionization detector.

in a large insulated box which can be cooled down to 260 K. After changing the setting of the recirculating thermostat by an increment of 10 K it takes about 1 h until the temperature of the diffusion bridge has equilibrated.

Pure carrier gas is flown through one of the flow tubes, while a constant trace gas concentration c_0 is maintained in the other. A concentration gradient is established along the capillaries. This gives rise to a constant flux J_D by molecular diffusion through the diffusion bridge described by Fick's first law of diffusion

$$J_D = -D \left(\frac{\partial c}{\partial z} \right)_t = D \left(\frac{c_0 - c}{l} \right) \quad (16)$$

where c is the trace gas concentration at the low concentration end. Pressure differences between the flow tubes are carefully eliminated to suppress trace gas transport by viscous flow through the capillaries. This requires that both flow tubes are totally symmetric. The difference pressure is monitored using a differential high accuracy pressure transducer (MKS model 398, measuring range 10^{-4} Torr to 1 Torr). By measuring the concentration ratio in both flow tubes, downstream the diffusion

bridge the diffusion coefficient is obtained by

$$D = \frac{\dot{V}l}{n\pi r^2} \frac{c}{c_0} \quad (17)$$

125 when $c \ll c_0$ where \dot{V} is the volume flow rate of the carrier gas. The ratio of mass transport by viscous flow to diffusion flow through the capillaries is given by (viscosity η)

$$\frac{J_V}{J_D} \approx \frac{r^2 \Delta p}{16\eta D}. \quad (18)$$

Therefore, the ratio of interfering viscous trace gas to mass flow by diffusion was minimized by using narrow capillaries. For the diffusion of NO_2 in N_2 at standard pressure and temperature, the fraction of viscous flow can be held $< 1\%$ when keeping the differential pressure $\Delta p < 2 \times 10^{-4}$ Torr. During the TT-experiments the differential pressure was maintained low so that
130 the fraction of viscous flow was less than 0.3%.

A trace gas detector is required which is linear over a wide concentration range and stable over time. Depending on the trace gas and the detector properties, the trace gas can either be detected by continuous mode (Fig. 3) or by peak mode (Fig. 4). The low concentration is determined with a random error of about 1%. The signal of the trace gas detector is fed into a A/D-converter with 16-bit resolution (Data Translation DT2705/5715A).

135 The supplier of the capillary columns used in this work as raw material for assembling the diffusion bridge reports the inner diameter with a systematic error of 10%. This is too much for the measurement of diffusion coefficients. Therefore, two segments of the column were used to determine the inner diameter by weighing an empty and a water filled section of the capillary column. Thereby the radius of the diffusion bridge capillaries was determined with a systematic error of 1%. We tried to validate the result by electron micrography of two cross sections of the column material. However, the systematic error of
140 the inner diameter measured by electron micrography is about 5%. When assuming a systematic error of the flow rate of 1.5%, this results in a total systematic error of the method of $< 4\%$. The random error of the method depends on the random error of about 1% of the determination of the lower concentration c .

Later on during the experiments it was found that some capillaries of the diffusion bridge became blocked by dust or condensed matter. Fortunately, the TT method can be utilized to measure diffusion coefficients of several species simultaneously
145 when using the peak mode and a gas chromatograph as detector. If the diffusion coefficient of one of the trace gases has been determined with another reliable technique at one temperature, this diffusion coefficient can be used as internal standard, see Fig. 5. It is assumed that the effective area of the capillaries is independent of temperature.

3 Results and discussion

To determine D_0 and b , the diffusion coefficients obtained at different temperatures were fitted by nonlinear regression to
150 Eq. (5). They were weighted by the inverse of their statistical error where available. The results are summarized in Table 3 together with the diffusion coefficients calculated by the Chapman-Enskog theory using Eq. (1) and the Fuller method using Eq. (10). The input parameters used are summarized in Table 1.

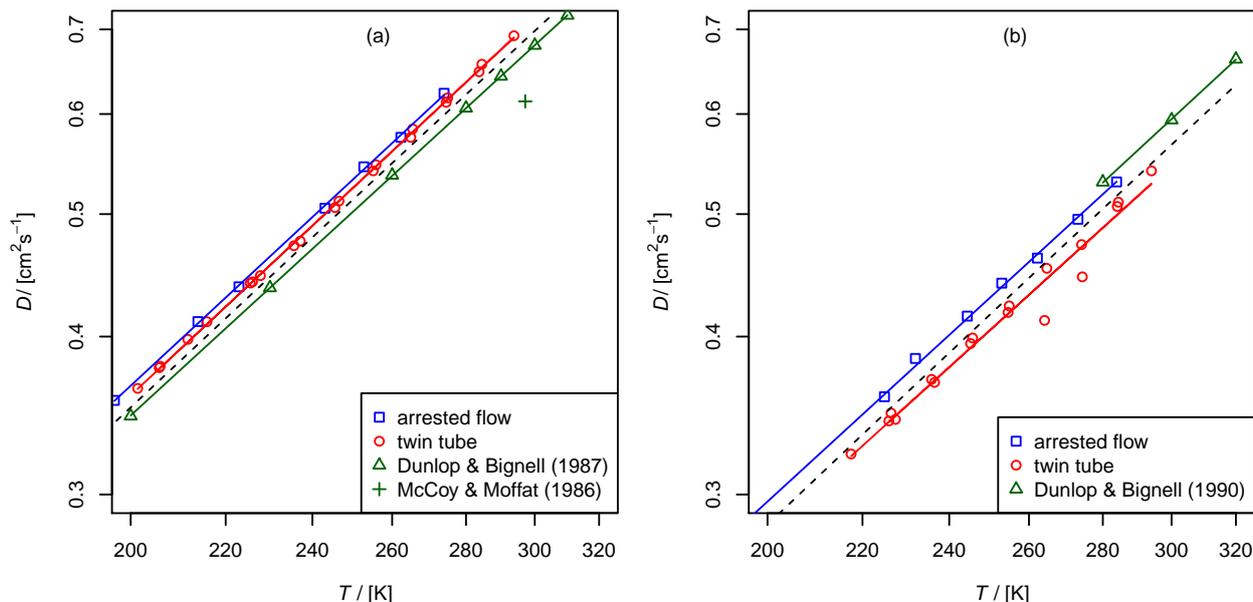


Figure 6. Comparison of diffusion coefficients obtained by the AF and TT method: (a) diffusion coefficient of methane in helium, compared to data of Dunlop and Bignell (1987) (fit $D_0 = 0.582 \text{ cm}^2\text{s}^{-1}$, $b = 1.66$) and $D = 0.614 \text{ cm}^2\text{s}^{-1}$ at 297 K of McCoy and Moffat (1986). (b) Diffusion coefficient of ethene in helium compared to data of Dunlop and Bignell (1990) (fit $D_0 = 0.508 \text{ cm}^2\text{s}^{-1}$, $b = 1.68$). The dashed line is calculated by the Lennard-Jones model.

3.1 Method validation

3.1.1 Diffusion of methane (CH_4) and ethene (C_2H_4) in helium and air

155 These gases were investigated with both methods for validation purposes. They are stable and non adsorbing, reference literature data of diffusion coefficients do exist. The diffusion coefficients in helium have been measured previously over a wider temperature range with high precision and accuracy by Dunlop and Bignell (1987) for methane and Dunlop and Bignell (1990) and ethene. Evaluated diffusion coefficients of hydrocarbons in air at 298 K are reported in the review of Tang et al. (2015). In addition, the diffusion coefficients can be calculated using the Lennard-Jones model and the Chapman-Enskog theory using
 160 Eq. (1).

We have used a flame ionization detector (Carlo Erba FID 40 with EL980 control unit) which is fast, sensitive and linear over a wide concentration range to measure the hydrocarbons. For the AF experiments, 0.1% of methane or ethane in helium or air were injected as a 300 ms pulse. About 20–26 sccm were used as flow rate. The arrest time was varied from 0 to 200 s. For the TT experiments 0.5% and 1% methane in air, 0.4% methane in He, 0.4% ethene in He and 0.5% ethene in air were admitted in
 165 the flow tube. Downstream of the diffusion bridge, the trace gas was analyzed using the peak mode setup, see Fig. 4.

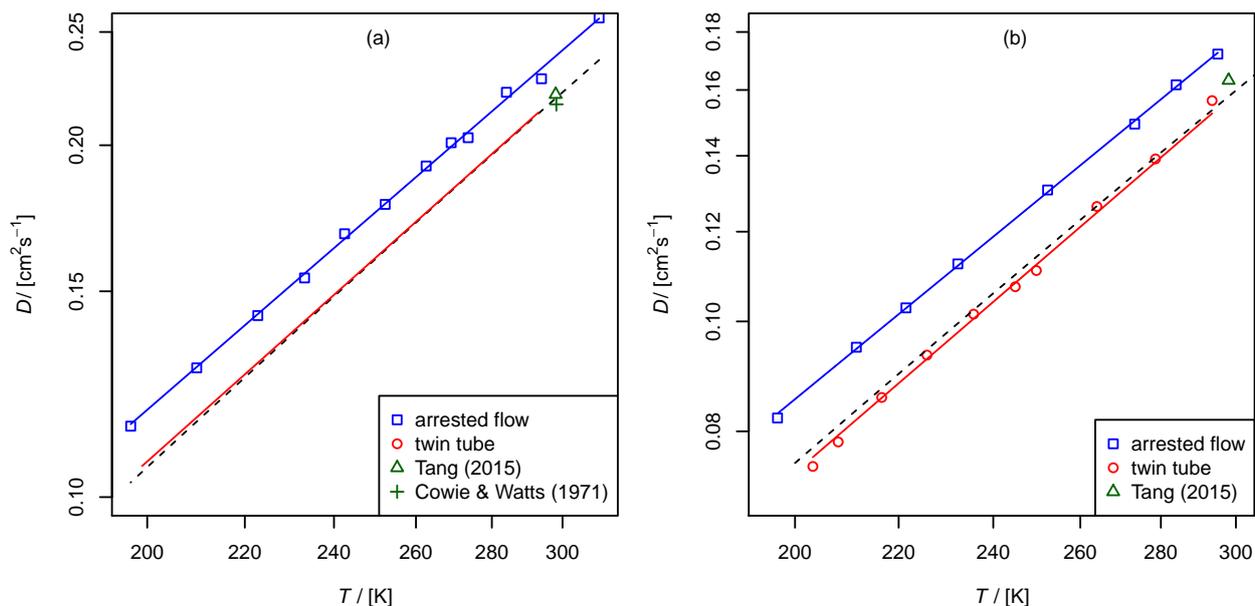


Figure 7. (a) Diffusion coefficient of methane in air, recommended value of $D = 0.221 \text{ cm}^2\text{s}^{-1}$ at 298 K of Tang et al. (2015) and $D = 0.217 \text{ cm}^2\text{s}^{-1}$ of Cowie and Watts (1971) at 298.2 K. For the TT method only the fit function is shown. (b) Diffusion coefficient of ethene in air as function of temperature, recommended value of Tang et al. (2015) $D = 0.163 \text{ cm}^2\text{s}^{-1}$ at 298 K. The dashed line is calculated using the Lennard-Jones model.

Table 2. Measured diffusion coefficients $D_{\sigma}D$ at the reference temperature T and [standard pressure \(101325 Pa\)](#) and reference data of diffusion coefficients D_r . With 298 K as reference temperature, the measured diffusion coefficients were extrapolated to this temperature using Eq. (5).

Species	Carrier	T [K]	D_r [cm^2s^{-1}]	$D/D_r - 1$	
				AF	TT
NO	He	273	0.624_a	6%	7%
CH ₄	He	273	0.582_b	6%	5%
C ₂ H ₄	He	273	0.508_c	-2%	-8%
CH ₄	air	298	0.221_d	8%	0%
C ₂ H ₄	air	298	0.163_d	8%	-4%

Source: a Dunlop and Bignell (1992), b Dunlop and Bignell (1987), c Dunlop and Bignell (1990), d Tang et al. (2015)

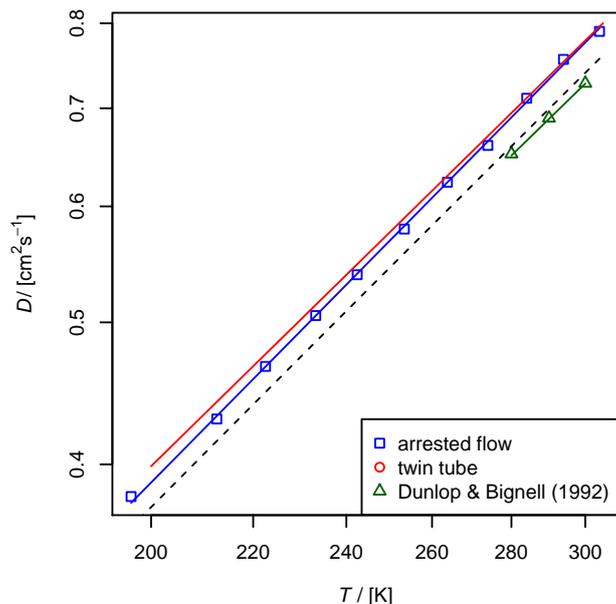


Figure 8. Diffusion coefficient of NO in He measured by AF and TT method as function of temperature. For the TT method only the fit function is shown. Experimental results of Dunlop and Bignell (1992) and a fit to Eq. (5) with $D_0 = 0.624$ and $b = 1.64$ are displayed as reference. The dashed line is calculated using the Lennard-Jones model.

The results are summarized in Table 2 and Table 3, in Fig. 6 and Fig. 7. Higher diffusion coefficients were found for the AF method compared to the TT-method.

3.1.2 Diffusion of nitric oxide (NO) in helium

NO was monitored by a chemiluminescence detector (Marić et al., 1989) which was adapted to the lower flow rates of the diffusion experiments. In the detector, NO reacts with O_3 in a low pressure reaction chamber (0.9–2 mbar) in a chemiluminescence reaction:



The emitted photons were detected using a Hamamatsu R562 photomultiplier tube.

For the AF experiment 100 ppm NO were injected as 300 ms pulse into the flow tube with a flow rate of 22.5 sccm. The valves were connected using stainless steel tubes. For the TT experiment 30–70 ppm NO in He were admitted into the diffusion bridge. The setup displayed in Fig 3 was used to monitor NO in the continuous mode. It was found that after measuring the high concentration c_0 , it needs several hours until the baseline has stabilized when measuring clean carrier gas. Therefore, it is not possible to measure c and c_0 in succession. Thus, first c_0 was measured at room temperature. Then the detector was switched to clean carrier gas until the baseline has stabilized. Then c was measured after lowering the temperature until the

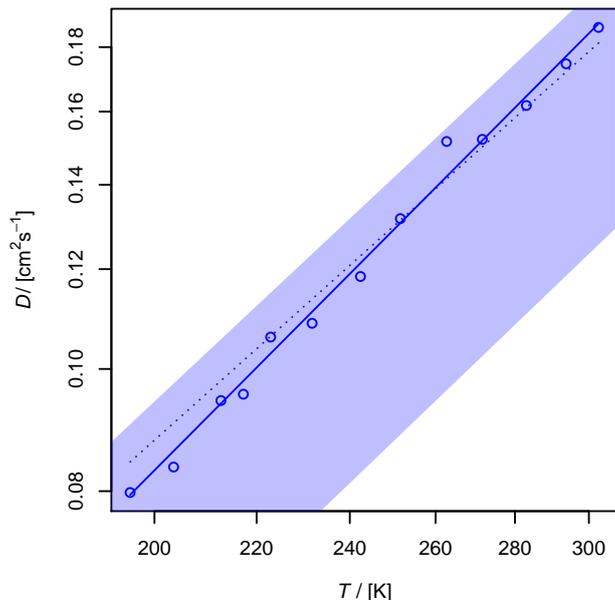


Figure 9. Diffusion coefficient of O_3 in air as function of temperature. The shaded area shows the diffusion coefficient expected by the Lennard-Jones model. The upper border is calculated with $D(T, 0.8\sigma_{AB}, 0.8\epsilon_{AB})$, the lower border with $D(T, 1.2\sigma_{AB}, 1.2\epsilon_{AB})$ using Eq. (1) corresponding to a 20% uncertainty of σ_{AB} and ϵ_{AB} . The dotted line is calculated by Fuller's model.

180 signal has stabilized. After arriving at 200 K, c_0 was measured again. Then the measurement was repeated by stepwise raising the temperature. Therefore, the complete measurement extended over several days.

The diffusion coefficients for NO obtained by the two methods are in fair agreement with the reference data of Dunlop and Bignell (1992), see Fig. 8 and Table 2. In contrast to diffusion coefficients of methane and ethane, the diffusion coefficients obtained by the TT method are slightly larger than the diffusion coefficients obtained by the AF method.

185 Comparing all diffusion coefficients obtained for stable gases to reference data, it is found that the deviation of the AF and TT method is less than 8% compared to the reference data. However, for the TT method this is a little more, as expected as theoretical systematic error, which can be explained by decreasing effective areas of the diffusion capillaries.

3.2 Diffusion of atmospheric trace gases

3.2.1 Diffusion of ozone (O_3) in air

190 The diffusion coefficient of ozone in air has never been measured before. Ivanov et al. (2007) reported $D = (0.53 \pm 0.03) \text{ cm}^2\text{s}^{-1}$ for the diffusion of O_3 in He at 298 K. Since ozone is an unstable but non-adsorbing species, only the AF method was used for the determination of the diffusion coefficient. A fast and sensitive ozone detector is required to record the ozone peaks leaving the column. A suitable detection technique is chemiluminescence arising from the reaction of ozone with Coumarin

47 (Lambda Physik, 7-diethylamino-4-methylcoumarin), adsorbed on silica-gel plates (Schurath et al., 1991). Chemilumines-
 195 cence is emitted in the range $\lambda = 440 - 550$ nm which is detected by photomultiplier (Hamamatsu 931 B). The anode current
 was admitted through a $100 \text{ k}\Omega$ resistor and measured as voltage by a microvoltmeter (Keithley model 155).

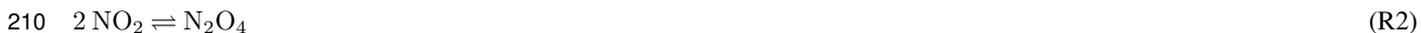
Ozone containing air was generated in an aluminum block enclosing an elliptically shaped polished chamber. A rod-shaped
 low-pressure Hg UV-lamp and a quartz tube with air running through are mounted parallel in the focal lines of the elliptically
 shaped chamber. Thereby, the UV radiation is focused to the air flowing through the quartz tube (Becker et al., 1975). 5 blocks
 200 arranged in series were used yielding an ozone concentration of about 40 ppm.

The injection time was varied from 250 ms to 500 ms and the arrest time was varied from 0 s to 360 s. It was found that the
 maxima of the eluted peaks did not coincide: peaks arrested longer were eluted later. Later it was found that this was caused
 by leaking sealings of the solenoid valves, which were made of neoprene. Therefore, at some temperatures only arrest times of
 less than 100 s were included in the fit of ζ_z^2 vs. t_a . The statistical error of the slope of ζ_z^2 vs. t_a was up to 3.7%.

205 With regard to the systematic error of 7% of the AF method for the diffusion coefficients (Section 3.1.1), the obtained value
 with error ranges is $D_0 = 0.15 \pm 0.01 \text{ cm}^2\text{s}^{-1}$. This value is in accordance with the value of $D_0 = 0.1444 \text{ cm}^2\text{s}^{-1}$ estimated
 by Massman (1998) from critical constants using the model of Chen and Othmer (1962).

3.2.2 Diffusion of nitrogen dioxide (NO_2) and dinitrogen tetroxide (N_2O_4) in helium and nitrogen

NO_2 is in equilibrium with its dimer



Therefore, a pure sample of NO_2 for determinations of D is not available. Chambers and Sherwood (1937) assumed that
 the ratio of $D(\text{NO}_2)/D(\text{N}_2\text{O}_4) = 1.43$ yielding $D_0(\text{NO}_2) = 0.121 \text{ cm}^2\text{s}^{-1}$ in nitrogen from their value of $D_0(\text{N}_2\text{O}_4) =$
 $(0.0845 \pm 0.0005) \text{ cm}^2\text{s}^{-1}$. Massman (1998) estimated $D_0(\text{NO}_2) = 0.146 \text{ cm}^2\text{s}^{-1}$ from $D_0(\text{N}_2\text{O}_4) = 0.101 \text{ cm}^2\text{s}^{-1}$ in ni-
 trogen reported by Sviridenko et al. (1973). Since NO_2 is an adsorbing species, the diffusion coefficient can only be measured
 215 using the TT method. The total flux of the pseudo species $\text{N}_{\text{IV}} = \text{NO}_2 + 2 \text{N}_2\text{O}_4$ though the capillaries is given by

$$J(\text{N}_{\text{IV}}) = \frac{D(\text{NO}_2)c_0(\text{NO}_2) + 2D(\text{N}_2\text{O}_4)c_0(\text{N}_2\text{O}_4)}{l}. \quad (19)$$

At higher temperatures and when keeping the concentration of NO_2 low, diffusion of N_2O_4 can be neglected. The degree of
 dissociation

$$\alpha = \frac{p(\text{NO}_2)}{p(\text{N}_{\text{IV}})} \quad (20)$$

220 can be calculated from the equilibrium constant

$$K = \frac{p^2(\text{NO}_2)}{p^\ominus p(\text{N}_2\text{O}_4)} = \frac{p(\text{N}_{\text{IV}})}{p^\ominus} \frac{2\alpha^2}{1 - \alpha} \quad (21)$$

where $p^\ominus = 1 \text{ bar}$. The equilibrium constant close to 250 K is estimated from JANAF Thermochemical Tables (NIST, 1998)

$$\ln K = 21.16 - 6878.1 \text{ K}/T. \quad (22)$$

Table 3. Results of Diffusion coefficient D_0 at standard temperature (273.15 K) and pressure (101325 Pa) and temperature coefficient b as results from the fit of Eq. (5) to the measured data. The errors listed are the errors obtained by the fit. Diffusion coefficients calculated by the Lennard-Jones and Fuller method are displayed for comparison. $D_0(\text{N}_2\text{O}_4)$ is estimated by nonlinear regression of Eq. (23) arbitrarily setting $b = 1.75$. For N_2O_5 the fit parameters of Eq. (24) are displayed.

Species	Carrier	Method	Experimental			Lennard-Jones model		Fuller et al. (1966)
			T [K]	D_0 [cm ² s ⁻¹]	b	D_0 [cm ² s ⁻¹]	b	D_0 [cm ² s ⁻¹]
CH ₄	He	AF	197–274	0.618 ± 0.002	1.68 ± 0.02	0.596	1.68	0.549
CH ₄	He	TT	201–294	0.610 ± 0.001	1.69 ± 0.01	"	"	"
CH ₄	air	AF	197–311	0.205 ± 0.001	1.75 ± 0.02	0.188	1.80	0.180
CH ₄	air	TT	199–293	0.188 ± 0.003	1.80 ± 0.04	"	"	"
C ₂ H ₄	He	AF	197–284	0.497 ± 0.002	1.66 ± 0.03	0.484	1.70	0.418
C ₂ H ₄	He	TT	218–294	0.468 ± 0.004	1.65 ± 0.08	"	"	"
C ₂ H ₄	air	AF	197–295	0.150 ± 0.000	1.81 ± 0.01	0.135	1.84	0.128
C ₂ H ₄	air	TT	203–294	0.133 ± 0.001	1.90 ± 0.05	"	"	"
NO	He	AF	196–304	0.662 ± 0.001	1.71 ± 0.02	0.633	1.68	0.757
NO	He	TT	200–305	0.667 ± 0.013	1.65 ± 0.02	"	"	"
NO ₂	He	TT	251–299	0.520 ± 0.001	1.93 ± 0.03	0.537	1.70	0.613
NO ₂	N ₂	TT	251–319	0.145 ± 0.001	1.94 ± 0.06	0.135	1.82	0.162
O ₃	air	AF	196–303	0.153 ± 0.001	1.97 ± 0.06	0.131	1.83	0.152
N ₂ O ₄	He	TT	202–299	0.221 ± 0.014	1.75 ± 0.25	0.388	1.71	0.440
N ₂ O ₄	N ₂	TT	204–319	0.084 ± 0.004	1.75 ± 0.25	0.090	1.88	0.115
ClONO ₂	He	AF	236–304	0.310 ± 0.004	1.57 ± 0.17	0.402	1.72	0.387
ClONO ₂	N ₂	AF	234–298	0.085 ± 0.001	2.27 ± 0.21	0.092	1.88	0.103
N ₂ O ₅	He	AF	245–298	0.300 ± 0.012	1.39 ± 0.23	0.381	1.73	0.405
N ₂ O ₅	N ₂	AF	246–298	0.081 ± 0.005	1.66 ± 0.26	0.085	1.91	0.106

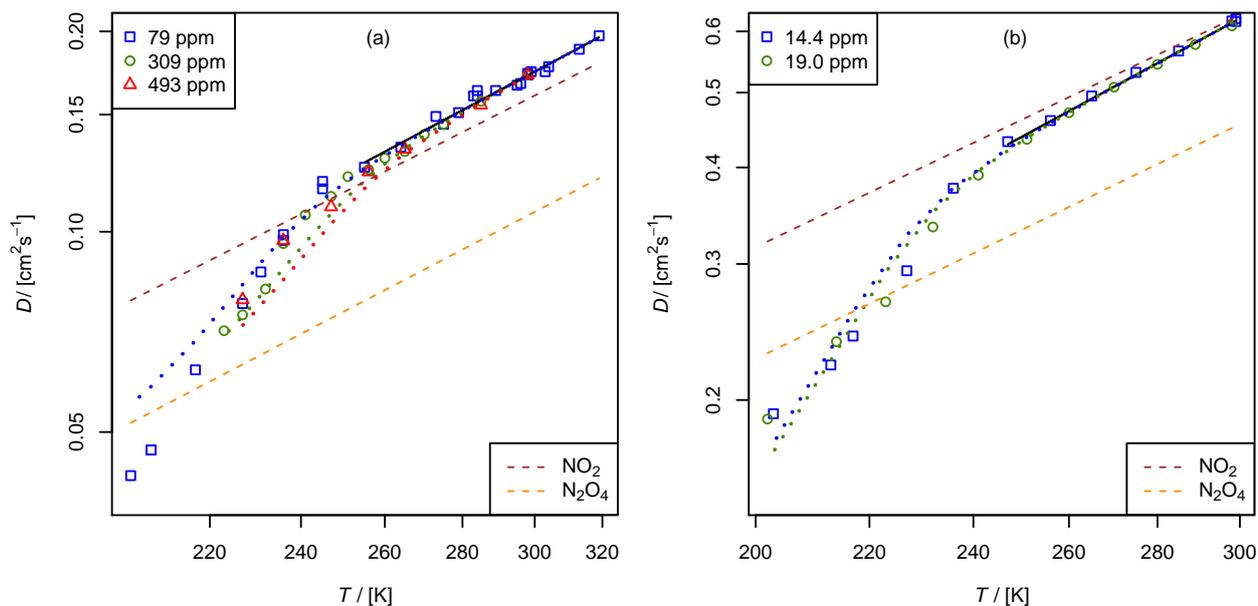


Figure 10. Apparent diffusion coefficient $D(N_{IV})$ of $\text{NO}_2/\text{N}_2\text{O}_4$ in N_2 (a) and He (b) as function of temperature. Black solid line: fit of Eq. (5) to data points with $\alpha > 0.95$. The dashed lines are calculated using the Lennard-Jones model. The dotted curves are fit curves of Eq. (23) by setting $b(\text{N}_2\text{O}_4) = 1.75$.

Actually $D(N_{IV})$ is determined using the TT experiment, depending on $D(\text{NO}_2)$ and $D(\text{N}_2\text{O}_4)$ (see Supplement S1):

$$225 \quad D(N_{IV}) = \alpha D(\text{NO}_2) + (1 - \alpha) D(\text{N}_2\text{O}_4) \quad (23)$$

A thermostatted permeation tube consisting of a PTFE tube (2.5 cm length, 4.3 mm diameter) closed with two Swagelok connectors filled with liquid N_2O_4 is used as NO_2 -source. The measurements were performed in the temperature range from 200 K to 300 K. NO_2 was measured as NO using the chemiluminescence analyzer preceded by a thermal converter. The converter, which consisted of a gold wire in a thin quartz tube heated to 540 K, was run with 1.5% – 3% methanol vapor instead of the more commonly used CO reagent (Langenberg et al., 1998). This eliminated poisoning of the gold wire which occurs when metal carbonyl impurities are present in the CO gas. The total conversion rate is unknown. However, for the TT experiment, it is only required that the conversion rate is independent of concentration, which was checked for the concentration range of 3 to 100 ppm NO_2 by a dynamic dilution experiment. NO_2 was monitored using the continuous mode of Fig. 5. In order to compensate for aging effects of the capillary bridge, CH_4 was admixed as an internal standard to the carrier gas flowing through the NO_2 permeation source. The 4-port valve in front of the detector was replaced by a 6-port valve (Valco UC10W, 125 μm sample loop) which enables detecting CH_4 by the flame ionization detector. For CH_4 in He $D_0 = (0.582 \pm 0.003) \text{ cm}^2\text{s}^{-1}$ at 273.15 K (Dunlop and Bignell, 1987) and for CH_4 in N_2 $D = (0.216 \pm 0.001) \text{ cm}^2\text{s}^{-1}$ at 298 K (Mueller and Cahill, 1964) was chosen as reference diffusion coefficient for the internal standard. The measurement was performed in

a similar manner like the NO measurement with the TT experiment by stepwise lowering and rising the temperature. Above
240 250 K the statistical error of the low concentration is about 3%. However, below 250 K signal stability was low. In contrast
to the measurements above 250 K, apparent $D(N_{IV})$ measured at a certain temperature by stepwise lowering the temperature
does not reproduce the value of $D(N_{IV})$ measured by stepwise rising the temperature.

Figure 10 displays the obtained diffusion coefficients of the pseudo species N_{IV} as function of temperature. It is obvious
that below 250 K the plot is deviating. Above 250 K diffusion of N_2O_4 can be neglected because N_2O_4 is mostly dissociated
245 in the concentration range of our study. To estimate the diffusion coefficient of NO_2 only data points with dissociation degree
 $\alpha > 0.95$ were included in the fit of Eq. (5). Thus, regarding the errors of the diffusion coefficients of the internal standards,
the diffusion coefficients for NO_2 at STP are $D_0 = (0.520 \pm 0.004) \text{ cm}^2\text{s}^{-1}$ and $D_0 = (0.145 \pm 0.002) \text{ cm}^2\text{s}^{-1}$ in helium and
nitrogen, respectively.

To determine $D_0(N_2O_4)$, Eq. (23) was fitted to experimental data of $D(N_{IV})$ vs. T and $p(N_{IV})$ by nonlinear regression.
250 The temperature dependency of $D(N_{IV})$ was described by Eq. (5) and α as function of temperature and $p(N_{IV})$ was calculated
using Eq. (21). $D_0(NO_2)$ and $b(NO_2)$ were taken as fixed input parameters from the fit of Eq. (5) as described above. However,
an independent determination of $D_0(N_2O_4)$ and $b(N_2O_4)$ was not possible. Therefore, $b(N_2O_4) = 1.75$ was set arbitrarily,
yielding the diffusion coefficients listed in Table 3. Since $b(N_2O_4)$ is expected in the range 1.5 to 2, the fit was repeated setting
 $b = 2$ and $b(N_2O_4) = 1.5$ to estimate the upper and lower limit of $D_0(N_2O_4)$ listed in Table 3. Compared to the diffusion
255 coefficient of N_2O_4 in N_2 , the diffusion coefficient in He determined by the fit is much lower than the diffusion coefficient
calculated by the Lennard-Jones model. In addition, our values are lower than the values of Sviridenko et al. (1973). We
therefore consider our diffusion coefficients for N_2O_4 to be unreliable. However, we can explain the observed temperature
dependency of $D(N_{IV})$ in the transition to lower temperatures.

3.2.3 Diffusion of chlorine nitrate ($ClONO_2$) in helium and nitrogen

260 Chlorine nitrate is an unstable compound. Therefore, the diffusion coefficient can only be measured using the AF-method and
not by the TT-method. Chlorine nitrate was prepared by the reaction (Davidson et al., 1987):



Cl_2O was prepared by admitting chlorine into a column filled with Raschig rings covered with freshly precipitated HgO
(Schmeisser et al., 1967):



The formed Cl_2O was condensed over N_2O_5 in a cold trap cooled with liquid nitrogen. Then the cold trap was cooled with
Ethanol at 193 K which was allowed to warm up to 253 K within about 15 h.

The identity of the product was checked by a FTIR spectrometer (Nicolet, model Protégé 460) with 10 m absorption path.
The spectrum was recorded with 1 cm^{-1} resolution. $ClONO_2$ was characterized by typical bands at 535–580 cm^{-1} , 750–
270 825 cm^{-1} , 1270–1320 cm^{-1} and 1695–1770 cm^{-1} compared to reference spectra measured by Davidson et al. (1987) and
Orphal et al. (1997). No contamination of N_2O_5 , NO_2 and HCl was found.

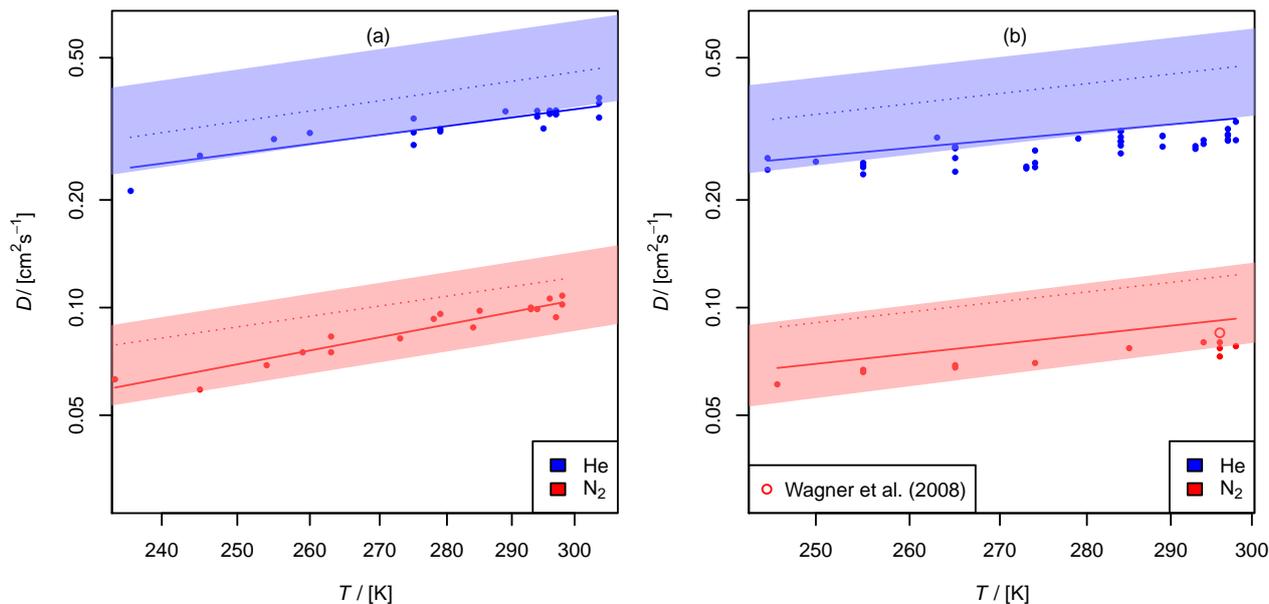


Figure 11. (a) Diffusion coefficient of chlorine nitrate in helium and nitrogen, (b) diffusion coefficient of dinitrogen pentoxide in helium and nitrogen as function of temperature. The experimental result of $0.085 \text{ cm}^2\text{s}^{-1}$ of Wagner et al. (2008) is displayed for comparison. The shaded areas show the diffusion coefficients expected by the Lennard-Jones model. The dotted lines are calculated by Fuller's model. The solid lines displayed for N_2O_5 are calculated using Eq. (24) extrapolating to $k_1 = 0 \text{ s}^{-1}$.

Chlorine nitrate dissociates by the equilibrium reaction:



The half-life of chlorine nitrate with respect to the thermal decomposition is 11 min at 300 K and about 7.8 h at 273 K. In addition to homogeneous dissociation, chlorine nitrate is lost by heterogeneous reaction with adsorbed water on the column surface (Tang et al., 2016):



To minimize this interference, the column was preconditioned with chlorine nitrate to remove moisture. During the first series of measurements with He as carrier gas, chlorine nitrate was continuously admitted into the column for 10–15 min prior to the experiments. However, due to desorption the base line stabilized only after some time. Therefore, during the second series of measurements with N_2 as carrier gas, series of peaks of chlorine nitrate were admitted into the column, until the peak size had stabilized. During one series of measurements 16 peaks covering arrest times from 0 to 100 s and a carrier gas flow rate of 19.6 sccm N_2 and 28.6 sccm He were measured. The diffusion coefficient was measured in a temperature range of 235 K to 300 K.

285 The detection of chlorine nitrate was performed as described by Anderson and Fahey (1990): an excess of NO (30–75 ppm) was added as constant flow of 6.5–8 sccm in N₂ before the detector by a T-tube. Behind the T-tube a glass capillary of 8 cm length and 0.1 cm inner diameter was mounted. The capillary was inserted in a stainless steel tube which was heated on a length of 1.7 cm to 433 K by two heating resistors. In the heating zone, chlorine nitrate is dissociated to ClO by Reaction R5. By subsequent scavenging reactions, NO is irreversibly removed:



295 A complete conversion of chlorine nitrate with NO is assumed. The drop of NO concentration equals the ClONO₂ concentration and was monitored using the chemiluminescence detector.

It is presumed that the chlorine nitrate loss processes during peak arrest is a pure first order process, which is a requirement for the application of the AF-method. To check the first order kinetics, logarithmic peak areas as measure for the chlorine nitrate concentration were plotted against the arrest time yielding a straight line. This validated the first order characteristic of the chlorine loss process. The first order loss constants ranged from $8.8 \times 10^{-4} \text{ s}^{-1}$ to $4.9 \times 10^{-3} \text{ s}^{-1}$. During the experiments with N₂ as carrier gas, loss rates increased with decreasing temperature. During the experiments with He, loss rates increased with increasing temperature. However, for the He experiments, a less effective preconditioning technique was applied, as described above.

300 The diffusion coefficients obtained at different temperatures are displayed in Fig.11 (a). When taking a systematic error of 7% for the AF method into account, the diffusion coefficients at STP are $D_0 = (0.31 \pm 0.03) \text{ cm}^2\text{s}^{-1}$ and $D_0 = (0.085 \pm 0.007) \text{ cm}^2\text{s}^{-1}$ in helium and nitrogen, respectively.

3.2.4 Diffusion of dinitrogen pentoxide (N₂O₅) in helium and nitrogen

Crystalline N₂O₅ was synthesized as described by Davidson et al. (1978) and Tang et al. (2014b): a small flow of pure NO is mixed with O₃/O₂ in a glass reactor and trapping the product at 193 K using a cold trap immersed in a cold ethanol bath. O₃ was generated in pure dry O₂ with a silent discharge ozone generator (Sorbios, model GSG). After mixing NO with O₃ / O₂ in the glass reactor, brown color appeared initially, indicating the formation of NO₂:





After about 3 h, addition of NO is stopped and the cooling bath of the cold trap is removed. To remove traces of NO₂, the product is transferred into a second cold trap using a O₃ / O₂ stream. By means of a cryostat, the synthesized N₂O₅ crystals were stored in an ethanol bath kept at 193 K. The identity of the product was checked by infrared spectroscopy. The spectrum was recorded with 1 cm⁻¹ resolution. N₂O₅ was characterized by typical bands at 750, 860, 1250, 1340 and 1725 cm⁻¹ reported by a reference spectrum measured by Cantrell et al. (1988).

The cold trap filled with N₂O₅ crystals was immersed in the bath of a cryostat (Lauda RLS 6) thermostatted at 235 K to 250 K. Dried carrier gas was admitted through the cold trap. An upper limit of the partial pressure of N₂O₅ in contact with the solid can be estimated from data of McDaniel et al. (1988). This results in an upper limit of the mole fraction of about 0.1% to 0.6%. The trace gas was admitted into the AF-experiments using short Teflon tubes.

N₂O₅ was detected as already described by Fahey et al. (1985): N₂O₅ is thermally decomposed to NO₂ and NO₃ radicals which are then titrated by NO to form NO₂



The drop in NO concentration is equal to the N₂O₅ concentration. NO was measured again by the chemiluminescence detector. Downstream of the AF-experiment, a constant flow of 6 – 8 sccm of 30 – 45 ppm NO was added.

The diffusion coefficient of N₂O₅ was measured in the temperature range 245 K to 298 K. Below 245 K no measurement was possible because N₂O₅ was totally adsorbed in the column. At one fixed temperature 16 peaks were measured using arrest times between 0 s and 100 s. At the beginning of a measurement series, the column was preconditioned with carrier gas containing N₂O₅ to remove moisture for 15 – 20 min. Prior of a measurement with arrest time, a peak without arrest time was pushed through the column. During one series of measurements 16 peaks covering arrest times from 0 to 100 s and a carrier gas flow rate of 19.3 sccm N₂ and 28.8 sccm He were measured.

Equation (5) was used to obtain $D_0 = (0.276 \pm 0.003) \text{ cm}^2\text{s}^{-1}$, $b = (1.0 \pm 0.2)$ in He and $D_0 = (0.0709 \pm 0.0006) \text{ cm}^2\text{s}^{-1}$, $b = (1.1 \pm 0.1)$ in N₂. Thus, the observed temperature coefficient is much lower than expected from Chapman-Enskog theory. As long as the N₂O₅ loss process is purely first order, the peak variance should not be affected by loss processes. To check this, peak areas were determined by integration. Plots of log (peak area) versus arrest times yielded apparent first order loss constants k_1 , ranging from $4 \times 10^{-3} \text{ s}^{-1}$ to $2 \times 10^{-2} \text{ s}^{-1}$. To check if D depends on k_1 , the diffusion coefficient was expressed by

$$D = D_0 \left(\frac{p_0}{p} \right) \left(\frac{T}{T_0} \right)^b \exp(ak_1) \quad (24)$$

as function of T and k_1 . It was found that D not only significantly depends on T , but also on k_1 . With P , the probability of the null hypothesis, it was found for He and N₂ $a = -(8 \pm 4) \text{ s}$ ($P < 0.05$) and $a = -(21 \pm 9) \text{ s}$ ($P < 0.06$), respectively. One reason for this may be that the order of the loss process of N₂O₅ is less than first order. The other fit parameters are displayed in Table 3. For He as carrier gas, a temperature coefficient of $b < 1.5$ was found. Due to the small temperature
350 range investigated, the temperature coefficient is rather uncertain. As final result, for He $D_0 = (0.30^{+0.03}_{-0.06}) \text{ cm}^2\text{s}^{-1}$ and N₂ $D_0 = (0.08^{+0.01}_{-0.02}) \text{ cm}^2\text{s}^{-1}$ are obtained when considering the systematic error of 7% for the AF method and the possible influence of dinitrogen pentoxide degradation on the diffusion coefficients obtained. Wagner et al. (2008) reported a diffusion coefficient of $0.085 \text{ cm}^2\text{s}^{-1}$ for N₂O₅ in N₂ at 760 Torr and 296 K, which is within the error limits of our result.

4 Conclusions

355 The AF method is best suited for the measurement of diffusion coefficients of volatile non adsorbing trace gases, even if the trace gas is unstable like ozone. However, it is required that the trace gas loss process is first order. Otherwise, the Gaussian peak shape is distorted and the variance of the peaks depends on species reactions. The TT method is best suited for stable but adsorbing species.

For stable non polar gases diffusion coefficients can be estimated from viscosity data using the Lennard-Jones model with
360 a systematic error of $< 5\%$, which is smaller than the systematic errors of less than 7% of the AF and TT-methods. For unstable atmospherically relevant trace gases and polar gases, the Lennard-Jones model parameters cannot be obtained by viscosity measurements. They can only be estimated from critical temperatures and volumes. Where dipole – induced dipole interactions come into play the systematic errors of the diffusion coefficients obtained in this way are in the same order as the errors of the diffusion coefficients of the unstable and reactive trace gases investigated in this study.

365 For the species investigated in this study, it is found that Fuller’s method overestimates diffusion coefficients of inorganic compounds with a systematic error of typically less than 35% and underestimates diffusion coefficients of organic compounds with a systematic error of less than 15%.

Code and data availability. Raw data of temperature dependent diffusion coefficients are included in the supplement. Sample code for the calculation of diffusion coefficients using the Lennard-Jones and the Fuller model written in the language R (R Core Team, 2017) is also
370 included.

Author contributions. Conceptualization and methodology (SS), investigation (TC, DH, SS), formal analysis and visualization (TC, DH, SL, SS), writing draft (SL), supervision and funding acquisition (US). All the authors have read and approved the final manuscript.

Competing interests. The authors declare that they have no conflict of interest.

Acknowledgements. This work was supported by the “Deutsche Forschungsgemeinschaft (DFG)” within the DFG priority program “Basics of the Impact of Air and Space Transportation on the Atmosphere”. TC acknowledges a doctoral grant from Karlsruher Institut für Technologie. We thank Ralf Rubrüder and Birgit Walter for setting up the AF experiment; Peter Boecker for the assistance measuring FT-IR spectra; Harald Saathoff for providing infrared spectra of N_2O_5 , HCl and NO_2 ; Ewald Hild for the preparation of electron micrographs of the fused silica columns and Dieter Gauer for technical assistance.

References

- 380 Anderson, L. C. and Fahey, D. W.: Studies with nitryl hypochlorite: thermal dissociation rate and catalytic conversion to nitric oxide using an NO/O₃ chemiluminescence detector, *The Journal of Physical Chemistry*, 94, 644–652, <https://doi.org/10.1021/j100365a027>, 1990.
- Becker, K., Heindrichs, A., and Schurath, U.: Ein transportables Gerät zur Kalibrierung von Ozonanalysatoren durch Messung der optischen Absorption, *Staub - Reinhalt. Luft*, 35, 326–329, 1975.
- Brokaw, R. S. and Svehla, R. A.: Viscosity and Thermal Conductivity of the N₂O₄ ⇌ 2 NO₂ System, *The Journal of Chemical Physics*, 44, 385 4643–4645, <https://doi.org/10.1063/1.1726692>, 1966.
- Cantrell, C. A., Davidson, J. A., McDaniel, A. H., Shetter, R. E., and Calvert, J. G.: Infrared absorption cross sections for N₂O₅, *Chem. Phys. Lett.*, 148, 358 – 363, [https://doi.org/10.1016/0009-2614\(88\)87288-9](https://doi.org/10.1016/0009-2614(88)87288-9), 1988.
- Chambers, F. S. and Sherwood, T. K.: Absorption of Nitrogen Dioxide by Aqueous Solutions, *Industrial & Engineering Chemistry*, 29, 1415–1422, <https://doi.org/10.1021/ie50336a022>, 1937.
- 390 Chen, N. and Othmer, D.: New generalized equation for gas diffusion coefficient, *J. Chem. Eng. Data*, 7, 37–41, <https://doi.org/10.1021/jc60012a011>, 1962.
- Cowie, M. and Watts, H.: Diffusion of Methane and Chloromethanes in Air, *Can. J. Chem.*, 49, 74–77, <https://doi.org/10.1139/v71-011>, 1971.
- Davidovits, P., Kolb, C. E., Williams, L. R., Jayne, J. T., and Worsnop, D. R.: Mass Accommodation and Chemical Reactions at Gas–Liquid 395 Interfaces, *Chem. Rev.*, 106, 1323–1354, <https://doi.org/10.1021/cr040366k>, PMID: 16608183, 2006.
- Davidson, J., Cantrell, C., Shetter, R., McDaniel, A., and Calvert, J.: Absolute Infrared Absorption Cross Sections for ClONO₂ at 296 and 223 K, *J. Geophys. Res.*, 92, 10 921–10 925, <https://doi.org/10.1029/JD092iD09p10921>, 1987.
- Davidson, J. A., Viggiano, A. A., Howard, C. J., Dotan, I., Fehsenfeld, F. C., Albritton, D. L., and Ferguson, E. E.: Rate constants for the reactions of O₂⁺, NO₂⁺, NO⁺, H₃O⁺, CO₃⁻, NO₂⁻, and halide ions with N₂O₅ at 300 K, *The Journal of Chemical Physics*, 68, 2085–2087, 400 <https://doi.org/10.1063/1.436032>, 1978.
- Dunlop, P. J. and Bignell, C.: Diffusion and thermal diffusion in binary mixtures of methane with noble gases and of argon with krypton, *Physica A*, 145, 584 – 596, [https://doi.org/10.1016/0378-4371\(87\)90009-4](https://doi.org/10.1016/0378-4371(87)90009-4), 1987.
- Dunlop, P. J. and Bignell, C. M.: The temperature and concentration dependencies of diffusion coefficients of some helium–hydrocarbon and helium–fluorocarbon systems, *The Journal of Chemical Physics*, 93, 2701–2703, <https://doi.org/10.1063/1.458908>, 1990.
- 405 Dunlop, P. J. and Bignell, C. M.: The Temperature and Concentration Dependences of Diffusion Coefficients of the Systems Ne–O₂, Kr–O₂, Xe–O₂ and He–NO, *Berichte der Bunsengesellschaft für physikalische Chemie*, 96, 1847–1848, <https://doi.org/10.1002/bbpc.19920961211>, 1992.
- Eller, A. S. D. and Sparks, J. P.: Predicting leaf-level fluxes of O₃ and NO₂: the relative roles of diffusion and biochemical processes, *Plant, Cell & Environment*, 29, 1742–1750, <https://doi.org/10.1111/j.1365-3040.2006.01546.x>, 2006.
- 410 Fahey, D., Eubank, C., Hübler, G., and Fehsenfeld, F.: A calibrated source of N₂O₅, *Atmospheric Environment (1967)*, 19, 1883 – 1890, [https://doi.org/10.1016/0004-6981\(85\)90013-7](https://doi.org/10.1016/0004-6981(85)90013-7), 1985.
- Fall, R. and Monson, R. K.: Isoprene Emission Rate and Intercellular Isoprene Concentration as Influenced by Stomatal Distribution and Conductance, *Plant Physiol.*, 100, 987–992, <https://doi.org/10.1104/pp.100.2.987>, 1992.
- Fuller, E. N., Schettler, P. D., and Giddings, J. C.: New Method for Prediction of Binary Gas-phase Diffusion Coefficients, *Industrial & 415 Engineering Chemistry*, 58, 18–27, <https://doi.org/10.1021/ie50677a007>, 1966.

- Gu, W., Cheng, P., and Tang, M.: Compilation and evaluation of gas phase diffusion coefficients of halogenated organic compounds, *R. Soc. Open Sci.*, 5, <https://doi.org/10.1098/rsos.171936>, 2018.
- Ivanov, A. V., Trakhtenberg, S., Bertram, A. K., Gershenson, Y. M., and Molina, M. J.: OH, HO₂, and Ozone Gaseous Diffusion Coefficients, *J. Phys. Chem. A*, 111, 1632–1637, <https://doi.org/10.1021/jp066558w>, PMID: 17298040, 2007.
- 420 Kirchner, W., Welter, F., Bongartz, A., Kames, J., Schweighoefer, S., and Schurath, U.: Trace gas exchange at the air/water interface: Measurements of mass accommodation coefficients, *J. Atmos. Chem.*, 10, 427–449, <https://doi.org/10.1007/BF00115784>, 1990.
- Knox, J. H. and McLaren, L.: A New Gas Chromatographic Method for Measuring Gaseous Diffusion Coefficients and Obstructive Factors., *Anal. Chem.*, 36, 1477–1482, <https://doi.org/10.1021/ac60214a017>, 1964.
- Kolb, C. E., Cox, R. A., Abbatt, J. P. D., Ammann, M., Davis, E. J., Donaldson, D. J., Garrett, B. C., George, C., Griffiths, P. T., Hanson, D. R., Kulmala, M., McFiggans, G., Pöschl, U., Riipinen, I., Rossi, M. J., Rudich, Y., Wagner, P. E., Winkler, P. M., Worsnop, D. R., and O’ Dowd, C. D.: An overview of current issues in the uptake of atmospheric trace gases by aerosols and clouds, *Atmos. Chem. Phys.*, 10, 10561–10605, <https://doi.org/10.5194/acp-10-10561-2010>, 2010.
- 425 Laisk, A., Kull, O., and Moldau, H.: Ozone Concentration in Leaf Intercellular Air Spaces Is Close to Zero, *Plant Physiol.*, 90, 1163–1167, <https://doi.org/10.1104/pp.90.3.1163>, 1989.
- 430 Langenberg, S., Proksch, V., and Schurath, U.: Solubilities and Diffusion of trace gases in cold sulfuric acid films, *Atmos. Environ.*, 32, 3129–3137, [https://doi.org/10.1016/S1352-2310\(97\)00490-1](https://doi.org/10.1016/S1352-2310(97)00490-1), 1998.
- Marić, D., Hans, W., and Schurath, U.: Measurements of N₂O photolysis coefficients in the stratosphere: Comparison with model calculations, *J. Atmos. Chem.*, 8, 19–40, <https://doi.org/10.1007/BF00053814>, 1989.
- Marrero, T. R. and Mason, E. A.: Gaseous Diffusion Coefficients, *J. Phys. Chem. Ref. Data*, 1, 3–118, <https://doi.org/10.1063/1.3253094>, 435 1972.
- Massman, W.: A review of the molecular diffusivities of H₂O, CO₂, CH₄, CO, O₃, SO₂, NH₃, N₂O, NO, and NO₂ in air, O₂ and N₂ near STP, *Atmos. Environ.*, 32, 1111 – 1127, [https://doi.org/10.1016/S1352-2310\(97\)00391-9](https://doi.org/10.1016/S1352-2310(97)00391-9), 1998.
- McCoy, B. J. and Moffat, A. J.: Arrested-flow Chromatographic Measurement of Gaseous Diffusion Coefficients, *Chem. Eng. Commun.*, 47, 219–224, <https://doi.org/10.1080/00986448608911765>, 1986.
- 440 McDaniel, A. H., Davidson, J. A., Cantrell, C. A., Shetter, R. E., and Calvert, J. G.: Enthalpies of formation of dinitrogen pentoxide and the nitrate free radical, *The Journal of Physical Chemistry*, 92, 4172–4175, <https://doi.org/10.1021/j100325a035>, 1988.
- Mueller, C. R. and Cahill, R. W.: Mass Spectrometric Measurement of Diffusion Coefficients, *The Journal of Chemical Physics*, 40, 651–654, <https://doi.org/10.1063/1.1725184>, 1964.
- Müller, B. and Heal, M. R.: The mass accommodation coefficient of ozone on an aqueous surface, *Phys. Chem. Chem. Phys.*, 4, 3365–3369, 445 <https://doi.org/10.1039/b202491h>, 2002.
- Neufeld, P. D., Janzen, A. R., and Aziz, R. A.: Empirical Equations to Calculate 16 of the Transport Collision Integrals $\Omega(l, s)^*$ for the Lennard-Jones (12–6) Potential, *The Journal of Chemical Physics*, 57, 1100–1102, <https://doi.org/10.1063/1.1678363>, 1972.
- NIST: JANAF Thermochemical Tables, in: NIST Standard Reference Database, vol. 13, <https://doi.org/10.18434/T42S31>, 1998.
- Orphal, J., Morillon-Chapey, M., Diallo, A., and Guelachvili, G.: High-Resolution Infrared Spectra and Harmonic Force Field of Chlorine Nitrate, *J. Phys. Chem. A*, 101, 1062–1067, <https://doi.org/10.1021/jp9619595>, 1997.
- 450 Patrick, R. and Golden, D. M.: Third-order rate constants of atmospheric importance, *Int. J. Chem. Kinet.*, 15, 1189–1227, <https://doi.org/10.1002/kin.550151107>, 1983.
- Poling, B. E., Prausnitz, J. M., and O’Connell, J. P.: *The Properties of Gases and Liquids*, McGraw-Hill, New York, 5 edn., 2004.

- R Core Team: R: A Language and Environment for Statistical Computing, R Foundation for Statistical Computing, Vienna, Austria, <http://www.R-project.org/>, (last access: February 18, 2020), 2017.
- 455 Schmeisser, M., Ruff, J. K., and Lustig, M.: Inorganic Syntheses, chap. Chlorine(1) Nitrate, pp. 127–130, Wiley-Blackwell, <https://doi.org/10.1002/9780470132401.ch34>, 1967.
- Schurath, U., Speuser, W., and Schmidt, R.: Principle and Application of a Fast Sensor for Atmospheric Ozone, *Fresenius J. Anal. Chem.*, 340, 544, <https://doi.org/10.1007/bf00322426>, 1991.
- 460 Sviridenko, Y. F., Makhin, V. A., and Shandorov, G. S.: Determining the diffusivity of nitrogen tetroxide, *Journal of engineering physics*, 24, 351–353, <https://doi.org/10.1007/bf00847664>, 1973.
- Tang, M., Keeble, J., Telford, P. J., Pope, F. D., Braesicke, P., Griffiths, P. T., Abraham, N. L., McGregor, J., Watson, I. M., Cox, R. A., Pyle, J. A., and Kalberer, M.: Heterogeneous reaction of ClONO₂ with TiO₂ and SiO₂ aerosol particles: implications for stratospheric particle injection for climate engineering, *Atmos. Chem. Phys.*, 16, 15397–15412, <https://doi.org/10.5194/acp-16-15397-2016>, 2016.
- 465 Tang, M. J., Cox, R. A., and Kalberer, M.: Compilation and evaluation of gas phase diffusion coefficients of reactive trace gases in the atmosphere: volume 1. Inorganic compounds, *Atmos. Chem. Phys.*, 14, 9233–9247, <https://doi.org/10.5194/acp-14-9233-2014>, 2014a.
- Tang, M. J., Telford, P. J., Pope, F. D., Rkiouak, L., Abraham, N. L., Archibald, A. T., Braesicke, P., Pyle, J. A., McGregor, J., Watson, I. M., Cox, R. A., and Kalberer, M.: Heterogeneous reaction of N₂O₅ with airborne TiO₂ particles and its implication for stratospheric particle injection, *Atmos. Chem. Phys.*, 14, 6035–6048, <https://doi.org/10.5194/acp-14-6035-2014>, 2014b.
- 470 Tang, M. J., Shiraiwa, M., Pöschl, U., Cox, R. A., and Kalberer, M.: Compilation and evaluation of gas phase diffusion coefficients of reactive trace gases in the atmosphere: Volume 2. Diffusivities of organic compounds, pressure-normalised mean free paths, and average Knudsen numbers for gas uptake calculations, *Atmos. Chem. Phys.*, 15, 5585–5598, <https://doi.org/10.5194/acp-15-5585-2015>, 2015.
- Taylor, G. I.: Dispersion of soluble matter in solvent flowing slowly through a tube, *Proceedings of the Royal Society of London. Series A. Mathematical and Physical Sciences*, 219, 186–203, <https://doi.org/10.1098/rspa.1953.0139>, 1953.
- 475 Taylor, G. I.: Conditions under which dispersion of a solute in a stream of solvent can be used to measure molecular diffusion, *Proceedings of the Royal Society of London. Series A. Mathematical and Physical Sciences*, 225, 473–477, <https://doi.org/10.1098/rspa.1954.0216>, 1954.
- Wagner, C., Hanisch, F., Holmes, N., de Coninck, H., Schuster, G., and Crowley, J. N.: The interaction of N₂O₅ with mineral dust: aerosol flow tube and Knudsen reactor studies, *Atmos. Chem. Phys.*, 8, 91–109, <https://doi.org/10.5194/acp-8-91-2008>, 2008.