

Response to Referee #1

We again thank the reviewer for the latest suggestions to improve the paper. The comments of the referee are repeated in bold letters while our response is given in normal text.

Line 144: Replace “added” with “applied”

Done, was changed as suggested.

Lines 216-217: Converting O3 VMR to [O3] does not involve the ideal gas law. I think what you want to say is something more like, “Air temperature and air pressure from SABER were used to calculate [M], [O2] (VMR of 0.21), and [N2] (VMR of 0.78) via the ideal gas law, and [M] was then used to convert SABER O3 VMR into [O3].”

Done, was changed as suggested.

Line 250: “level” should be “levels”

Done.

Line 589: Again, it’s my opinion that “above” and “below” are ambiguous when discussing profile values (can be confused with height instead of value). I’d suggest changing “always 10 % above or below O3 in chemical equilibrium” to something like “always within 10% of O3...” or “always 10% greater or lesser than O3...”

We agree with your suggestion but simply missed this error. Was changed to: “greater or lesser than...”.

Line 595-596: please explicitly state the criterion

Was rewritten as follows:”..., but only because the previously used criterion $(OH(9-7)+OH(8-6)VER > 10 \times G \times B)$ is valid.“

Response to Referee #2

We again thank the reviewer for the latest suggestions to improve the paper. The comments of the referee are repeated in bold letters while our response is given in normal text.

General Comments:

1. Kalogerakis et al. (2016) only reported that $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(1\text{D})$ is an important deactivation channel of $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{products}$. They did not provide any rate or branching ratio of the channel. To our understanding, “important” means “not negligible” but it does not mean “dominating”. Thus, we assumed that $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(1\text{D})$ has to occur but this channel is not necessarily the fastest deactivation path of $\text{OH}(v=9)+\text{O}(3\text{P})$.

This was stated in the text (l. 491-494):

“However, not much is known about the individual branching ratios of R11 except that $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(1\text{D})$ is an important deactivation channel but not necessarily the dominating one (Kalogerakis et al., 2016).”

...

Furthermore, we have to emphasize that, at least to our knowledge, the rates and deactivation schemes applied in the OH model are not in conflict with ANY laboratory measurements but partly disagree with other model studies.

Kalogerakis et al. (2016) measured the rate coefficient of the $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(1\text{D})$ pathway. Accounting for mesospheric temperatures, the authors suggested a rate of $(2.3\pm 1)e-10$ cm³/sec. Your results suggest, for the same pathway, a rate coefficient $0.2 * 2.3e-10$ cm³/sec, which is $4.6e-9$ cm³/sec, about 3 times lower than the lower bound suggested by Kalogerakis et al. (2016). Whether the reaction is the dominating one or not, the rate coefficient you use for this pathway still does not agree with laboratory results.

In general, many O retrieval studies which use OH models only need to consider $\text{OH}(v=9)$ and/or $\text{OH}(v=8)$ and therefore, state to state pathways coupling higher and lower levels do not need to be considered carefully. Because you are fitting four bands and require detailed description of $\text{OH}(v=3-9)$, these state to state pathways become very important. Currently your OH model disagrees with laboratory measurements, so a sentence in the conclusion and abstract should be added as a disclaimer noting the differences.

We thank the reviewer for highlighting this issue.

We extended the Abstract (l. 20-22):

“The results also provide general support of the recently proposed mechanism $\text{OH}(v)+\text{O}(3\text{P})\rightarrow\text{OH}(0\leq v'\leq v-5)+\text{O}(1\text{D})$ but suggest slower rates of $\text{OH}(v=8,7,6,5)+\text{O}(3\text{P})$, partly disagreeing with laboratory experiments.”,

changed Lines 494-495 to:

“However, not much is known about the individual branching ratios of R11 except that $\text{OH}(v=9)+\text{O}(3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(1\text{D})$ is an important deactivation channel but not necessarily the dominating one (Kalogerakis et al., 2016). These authors suggested a rate of $2.3(\pm 1.0)\times 10^{-10}$ cm³ s⁻¹ and noted that this rate might be slower due to the involvement of excited surfaces.”,

and added in the Conclusions (l. 647-649):

“In particular, the rate of $\text{OH}(v=9)+\text{O}(^3\text{P})\rightarrow\text{OH}(v=3)+\text{O}(^1\text{D})$ is about 3 times slower than the lower limit reported by Kalogerakis et al. (2016).”

2. "Figure 6 displays the vertical profiles of [O(3P)] and [H] obtained by the Best fit model in comparison with the results derived from SABER OH(9-7)+OH(8-6) VER only (Mlynczak et al., 2018). The [O(3P)] profiles seen in Fig. 6a agree below 85 km but the Best fit model shows gradually larger values in the altitude region above. These larger values are caused by the different deactivation rates and schemes of OH(v)+O(3P), agreeing with general pattern reported in Panka et al. (2018). We have to point out that other studies (e.g. von Savigny and Lednyts'kyi, 2013) observed a pronounced [O(3P)] maximum of about $8\times 10^{11}\text{ cm}^{-3}$ at 95 km."

It is clear in Figure 6 that your O begins to gradually increase in the upper altitudes. In a recent paper by Zhu & Kaufmann et al (2018), the authors show that O retrievals by Mlynczak et al. (2018), Panka et al. (2018), and Zhu and Kaufmann (2018) all show O densities which peak around 95 km (see Figures 2 and 3). I am confused by the statement where you report that your O follows the general pattern of Panka et al. (2018) as this increasing pattern is not shown in Panka et al. (2018) or Zhu and Kaufmann (2018). Please remove this sentence or clarify your meaning.

This statement focused on the altitude region below 95 km and referred to the general increase of [O(3P)] below 95 km.

Thus, we added (l. 531-533):

“Between 85 km and 95 km, these larger values are caused by the different deactivation rates and schemes of $\text{OH}(v)+\text{O}(^3\text{P})$, agreeing with general pattern reported in Panka et al. (2018).”

Model results of OH airglow considering four different wavelength regions to derive night-time atomic oxygen and atomic hydrogen in the mesopause region

Tilo Fytterer¹, Christian von Savigny², Martin Mlynczak³, Miriam Sinnhuber¹

5 ¹Institute for Meteorology and Climate Research, Karlsruhe Institute of Technology, Eggenstein-Leopoldshafen, 76344, Germany

²Institute of Physics, University of Greifswald, Greifswald, 17489, Germany

³NASA, Langley Research Center, Hampton, Virginia, 23681-2199, USA

Correspondence to: Miriam Sinnhuber (miriam.sinnhuber@kit.edu)

10 **Abstract.** Based on the zero dimensional box model CAABA/MECCA-3.72f, an OH airglow model was developed to derive night-time number densities of atomic oxygen ($[O(^3P)]$) and atomic hydrogen ($[H]$) in the mesopause region (~75-100 km). The profiles of $[O(^3P)]$ and $[H]$ were calculated from TIMED/SABER satellite OH airglow emissions measured at 2.0 μm . The two target species were used to initialize the OH airglow model, which was empirically adjusted to fit four different OH airglow
15 emissions observed by the satellite/instrument configuration TIMED/SABER at 2.0 μm and at 1.6 μm as well as measurements by ENVISAT/SCIAMACHY of the transitions OH(6-2) and OH(3-1). Comparisons between the “Best fit model” obtained here and the satellite measurements suggest that deactivation of vibrationally excited OH(v) via OH($v \geq 7$)+O₂ might favour relaxation to OH($v' \leq 5$)+O₂ by multi-quantum quenching. It is further indicated that the deactivation pathway to OH($v' = v - 5$)+O₂
20 dominates. The results also provide general support of the recently proposed mechanism OH(v)+O(³P)→OH($0 \leq v' \leq v - 5$)+O(¹D) but suggest slower rates of OH($v = 8, 7, 6, 5$)+O(³P), partly disagreeing with laboratory experiments. Additionally, deactivation to OH($v' = v - 5$)+O(¹D) might be preferred. The profiles of $[O(^3P)]$ and $[H]$ derived here are plausible between 80 km and 95 km but should be regarded as an upper limit. The values of $[O(^3P)]$ obtained in this study agree with the
25 corresponding TIMED/SABER values between 80 km and 85 km, but are larger from 85 to 95 km due to different relaxation assumptions of OH(v)+O(³P). The $[H]$ profile found here is generally larger than

TIMED/SABER [H] by about 50 % from 80 to 95 km, which is primarily attributed to our faster OH($v=8$)+O₂ rate.

1 Introduction

30 Atomic oxygen in its ground state (O(³P)) and atomic hydrogen (H) strongly influence the energy budget in the mesopause region (~75-100 km) during day and night (Mlynczak and Solomo, 1993), and consequently affect atmospheric air temperature, wind, and wave propagation (Andrews et al., 1987). Therefore, an improved knowledge of the abundance of O(³P) and H is of great importance when studying the mesopause region. At these altitudes, O(³P) has a direct impact on the heating rates by participating in several exothermic chemical reactions (Mlynczak and Solomon, 1993, their Table 4).
35 But O(³P) also contributes to radiative cooling by exciting CO₂ via collisions, leading to increased infrared emissions of CO₂ and partly opposing the O(³P) chemical heating effect. Night-time H plays a crucial role in the mesopause region due to the destruction of ozone (O₃) which is accompanied by the release of a considerable amount of heat (Mlynczak and Solomon, 1993). This chemical reaction
40 additionally leads to the production of vibrationally excited hydroxyl radicals (OH($v>0$)) up to the vibrational level $v=9$, causing the formation of OH emission layers in the atmosphere (Meinel bands; Meinel, 1950).

Direct measurements of O(³P) and H are relatively rare because as atomic species they do not have observable vibration-rotation spectra. Consequently, measuring these species in the mesopause region
45 by remote sensing requires complex methods while in situ observations are rather expensive (e.g. Mlynczak et al., 2004; Sharp and Kita, 1987). Thus, there exists no global data set based on direct observations. As a consequence, an indirect method was introduced by Good (1976) to derive [O(³P)] and [H] during night, using OH airglow emissions. This approach was also adapted by Mlynczak et al. (2013a; 2014; 2018) which derived a global data set of night-time [O(³P)] and [H] in the mesopause
50 region from satellite observations of OH(v). The method is based on the assumption of chemical steady state of O₃ and further depends on several radiative lifetimes, chemical reactions, and physical processes involving OH(v). However, the corresponding total rate coefficients and branching ratios are

still not sufficiently known, and thus present a large source of uncertainty in the derivation of $[O(^3P)]$ and $[H]$.

55 There are two major issues currently discussed in the literature which considerably affect the overall abundance of derived $O(^3P)$ and H . The first problem addresses the underlying deactivation schemes of $OH(v)$ from the higher excited state v to the lower excited state v' ($v' < v$) by collisions with O_2 . This can generally occur via sudden death ($OH(v)+O_2 \rightarrow OH(v'=0)+O_2$), single-quantum ($OH(v)+O_2 \rightarrow OH(v'=v-1)+O_2$), or multi-quantum ($OH(v)+O_2 \rightarrow OH(v' < v)+O_2$) quenching. However, 60 in case of the sudden death approach, it is still unknown where such a huge amount of excess energy is transferred. The second crucial point comprises the deactivation scheme and the total rate of $OH(v)+O(^3P)$, including the new pathway $OH(v)+O(^3P) \rightarrow OH(0 \leq v' \leq v-5)+O(^1D)$ suggested by Sharma et al. (2015).

Over the last three to four decades, several model studies attempted to fit OH airglow measurements, 65 using different rates and schemes for the deactivation of $OH(v)$ by O_2 and by $O(^3P)$. And at least to our knowledge, there is no general agreement about which model is correct. The deactivation of $OH(v)$ by O_2 in many models (e.g. von Savigny et al., 2012; Mlynczak et al., 2013a; Grygalashvyly et al., 2014; Panka et al., 2017) is based on the model proposed by Adler-Golden (1997). It assumes a combination of multi-quantum and single-quantum quenching and was derived from theoretical considerations and 70 ground-based observations. Xu et al. (2012) investigated measurements from the Sounding of the Atmosphere using Broadband Emission Radiometry (SABER) instrument on board the NASA Thermosphere-Ionosphere-Mesosphere Energetics and Dynamics (TIMED) satellite of the OH airglow emissions at $2.0 \mu m$ and at $1.6 \mu m$. Their results support the model of Adler-Golden (1997) but suggest slower total $OH(v)+O_2$ rates. They further exclude the sudden death mechanism as a possible 75 deactivation scheme. There are also two theoretical studies (Shalashilin et al., 1995; Caridade et al., 2002) which investigated $OH(v)$ deactivation via O_2 , both supporting a combination of multi-quantum and single-quantum quenching similar to the model of Adler-Golden (1997).

However, Russell and Lowe (2003) and Russell et al. (2005) analyzed $OH(8-3)$ and $O(^1S)$ airglow emissions measured by the Wind Imaging Interferometer (WINDII) instrument on board the Upper 80 Atmospheric Research Satellite (UARS). Both airglow emissions were used to derive separate data sets

of $[O(^3P)]$ and the best agreement between these two $[O(^3P)]$ data sets was obtained when a sudden death scheme for $OH(v)+O_2$ quenching was applied. Kaufmann et al. (2008) investigated several OH airglow spectra between 1 μm and 1.75 μm measured by the Scanning Imaging Absorption Spectrometer for Atmospheric Cartography (SCIAMACHY) instrument on board the Environmental Satellite (ENVISAT). They found best agreement between their model and the measured OH airglow spectra when a combination of sudden death and single-quantum quenching was used.

Vibrationally dependent rates of $OH(v)+O(^3P)$ were determined by Varandas (2004) and Caridade et al. (2013), using quasi-classical trajectory calculations. Their results suggest that deactivation occurs via a chemical reaction as well as multi-quantum quenching. Kalogerakis et al. (2011) obtained a deactivation rate of $OH(v=9)+O(^3P)$ from laboratory experiments which is several times larger than the rate from these calculations. But applying this fast quenching rate led to non-physical $[O(^3P)]$ values and associated heating rates (Smith et al., 2010; Mlynczak et al., 2013a). Thus, Sharma et al. (2015) proposed a new mechanism $OH(v)+O(^3P)\rightarrow OH(0\leq v'\leq v-5)+O(^1D)$ to account for results from both theory and experiment. Very recent laser experiments and model studies support this new pathway while the exact values of the branching ratios and total loss rates are still not known (Kalogerakis et al., 2016; Panka et al., 2017). However, recently published results by Mlynczak et al. (2018) oppose this mechanism. They also applied the new rate of Kalogerakis et al. (2011) for $OH(v=9)+O(^3P)$. But in order to get the annual energy budget into near balance, it was necessary to assume that at least $OH(v=9)+O(^3P)$ occurs via single-quantum relaxation. Additionally, the rate of $OH(v=8)+O_2$ had to be reduced and is considerably smaller than the value reported from Adler-Golden (1997).

The newly suggested rates of $OH(v)+O(^3P)$ were applied in different models to derive $[O(^3P)]$ in the mesopause region. Mlynczak et al. (2018) used SABER OH airglow emissions observed at 2.0 μm to derive $[O(^3P)]$ and assumed rates of $3.0\times 10^{-10}\text{ cm}^3\text{ s}^{-1}$ and $1.5\times 10^{-10}\text{ cm}^3\text{ s}^{-1}$ for $OH(v=9)+O(^3P)$ and $OH(v=8)+O(^3P)$, respectively. They further stated that deactivation of $OH(v=9)+O(^3P)$ has to occur via single-quantum quenching and that the $OH(v=8)+O_2$ rate has to be smaller than known from laboratory measurements to get global annual energy budget into near balance. Panka et al. (2018) simultaneously investigated SABER OH airglow emissions measured at 2.0 μm and 1.6 μm , while applying faster rates

for OH($v=8$)+O(3P) and OH(v)+O₂. Their [O(3P)] values agree within the corresponding errors with the results reported by Mlynczak et al. (2018) above ~87 km but are larger in the altitude region below. The authors also demonstrated the high sensitivity of the derived [O(3P)] from O(3P) quenching rates applied in their model. Zhu and Kaufmann (2018) analyzed SCIAMACHY OH(9-6) transition. They used a value of $2.3 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$ for OH($v=9$)+O(3P) which is lower than the one applied in the two previous studies, resulting in generally lower [O(3P)] values in the altitude region above 87 km. Their rate for OH($v=9$)+O₂ lies between the corresponding rates of the two other studies, and consequently their [O(3P)] is also between the [O(3P)] values of these two studies below 87 km. Thus, recent publications indicate that the rate of OH($v=9,8$)+O(3P) might be slower than previously suggested in Sharma et al. (2015). But this problem needs further attention because all three papers derive different [O(3P)], depending on the data sets investigated.

In order to address the two major issues stated above, this paper is focused on the development of a zero dimensional box model for atmospheric OH airglow with the intention to derive night-time [O(3P)] and [H] in the mesopause region. The model considers the formation of OH(v) via H+O₃ and deactivation of OH(v) due to spontaneous emission of photons, chemical reactions and physical collisions with atmospheric air compounds N₂, O₂, and O(3P). We used the indirect method introduced by Good (1976) and derived night-time [O(3P)] and [H] from TIMED/SABER OH emissions at ~2.0 μm , while also considering the OH airglow observations from TIMED/SABER at ~1.6 μm as well as the OH(6-2) and OH(3-1) transitions measured by ENVISAT/SCIAMACHY. Further sensitivity runs were carried out to estimate the uncertainty on the derived values of [O(3P)] and [H] due to the different deactivation schemes, overall rate constants, and branching ratios.

2 Data and method

2.1 Satellite measurements

2.1.1 ENVISAT/SCIAMACHY

The SCIAMACHY instrument (Bovensmann et al., 1999) was an 8-channel spectrometer on board

135 ENVISAT, providing atmospheric OH airglow emission measurements between ~220 nm and ~2380 nm. ENVISAT was launched into a polar and sun-synchronous orbit and crossed the equator at ~10 LT and ~22 LT. The ENVISAT mission started in March 2002 and SCIAMACHY was nearly continuously operating until the end of the mission in April 2012 caused by a spacecraft failure. The SCIAMACHY instrument performed measurements in different observations modes, including night-time (~22 LT) limb scans over the tangent altitude range ~75-150 km. These measurements are only available throughout the year at latitudes between the equator and 30° N.

140 In this paper, we used SCIAMACHY level 1b data v7.04 to retrieve OH airglow volume emission rates (VERs) of the OH(3-1) and OH(6-2) bands in the wavelength ranges of 1515-1546 nm and 837.5-848.0 nm, respectively. The retrieval approach applied here is very similar to the one described in von Savigny et al. (2012). The retrieval does not cover the complete spectra of the OH(3-1) and OH(6-2) bands, and consequently a “correction factor” of 2.48 for OH(3-1) VER and 2.54 for OH(6-2) VER was applied to
145 account for the entire band emissions at mesopause temperature. The data set further includes corrections for misalignments and other measurement errors (Gottwald et al., 2007). Investigations performed by Bramstedt et al. (2012) showed a drift of the SCIAMACHY tangent height of less than 20 m year⁻¹ which is negligible for our study.

The uncertainties of the OH(3-1) VER and OH(6-2) VER retrievals from SCIAMACHY limb
150 observations correspond to the propagated uncertainties of the observed limb emission rate (LER) profiles. The latter are estimated from the LER values in the tangent height range between 110 km and 150 km, where the actual atmospheric emissions should be zero. The VER uncertainties are first determined for daily and zonally averaged data. The uncertainties used in this analysis correspond to the mean uncertainties averaged over all days with co-located SCIAMACHY and SABER observations.

155 2.1.2 TIMED/SABER

The SABER instrument (Russell et al., 1999) on board the TIMED satellite has been nearly continuously operating since January 2002, collecting over 98 % of all possible data. The instrument scans the atmosphere from the surface up to altitudes of ~400 km while providing a vertical resolution of about 2 km throughout the entire height interval. Due to the geometry of the satellite orbit and the

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160 regular yaw manoeuvres every ~60-65 days, SABER only provides complete coverage of the latitude
range between ~55° S and ~55° N. The SABER instrument measures the OH VERs at ~2.0 μm and at
~1.6 μm which approximately corresponds to the transitions of OH(9-7)+OH(8-6) and OH(5-3)+OH(4-
2), respectively. The contribution of OH(7-5) to OH VER at 2.0 μm and of OH(3-1) to OH VER at 1.6
165 μm is only about a few percents (Xu et al., 2012; Mlynczak et al., 2013a) and is neglected in this paper.
In this study, we used the SABER Level 2A data v2.0 of the “unfiltered” OH VERs at 2.0 μm and at 1.6
μm, the air temperature and pressure, and the volume mixing ratios (VMRs) of O₃ (derived at 9.6 μm).
There are also SABER O₃ measurements at 1.27 μm but these observations are not available during
night. New night-time VMRs of O(³P) and H (Mlynczak et al., 2018) were used for comparison with the
results derived from our model. The “unfilter” factor applied to OH VER adjusts the originally
170 measured OH VER by the SABER instrument to the total VER emitted by OH in the corresponding
vibrational bands, while considering the shape, width, and transmission of the SABER broadband filters
(Mlynczak et al., 2005). Outliers were excluded by screening the data as suggested by Mlynczak et al.
(2013a). The SABER data used here were further restricted to observations between 21 LT and 23 LT to
approximately match the SCIAMACHY measurement time at ~22 LT. In order to be consistent with the
175 naming of the SCIAMACHY OH airglow observations, the SABER OH airglow at 2.0 μm and at 1.6
μm are referred to as OH(9-7)+OH(8-6) and as OH(5-3)+OH(4-2) throughout the paper.
The total uncertainty of SABER OH airglow data used here comprises three different error sources.
Since we used climatology of the measurements (see Sect. 2.2), there are sufficient samples that the
random noise component of the total uncertainty is essentially zero. The remaining two major terms are
180 the absolute calibration error (<5 %) and the “unfilter” factor error (<3 %). Assuming a root-sum-square
propagation of the individual uncertainties, this results in a total uncertainty of about 6 % for all data
points presented in this study.

2.2 Method

In order to minimize uncertainties between SABER and SCIAMACHY due to different measurement
185 characteristics, we focused on the latitude range from 0° to 10° N, which was covered by both
instruments throughout the entire year. A broader latitude band is not recommended because SABER

and SCIAMACHY do not uniformly cover the same latitudes, leading to disagreements between the real latitude of the observations and the nominal latitude of the interval. The accepted profiles of both instruments within the chosen latitude interval were averaged to zonal mean nightly mean values. All these zonal mean nightly means from January 2003 to December 2011 were used to calculate a climatology, including only days on which both SCIAMACHY and SABER data are available.

The approach to derive $[O(^3P)]$ and $[H]$ applied here was developed by Good (1976) and is described in detail in Mlynchak et al. (2013a). Thus, we only give a brief summary here. The measured SABER OH(9-7)+OH(8-6) VER (photons $cm^{-3} s^{-1}$) is given by Eq. (1):

$$OH(9-7)+OH(8-6) VER = k_1 [H][O_3]G(f_\nu, A_{\nu\nu'}, C_{\nu\nu'}), \quad (1)$$

where k_1 is the rate constant of the chemical reaction $H+O_3$, representing direct production. The function G (Eq. (2)) comprises all relevant production and loss processes of OH(9-7) VER and OH(8-6) VER:

$$G = \frac{f_9}{A_9 + C_9} A_{97} + \frac{f_8}{A_8 + C_8} A_{86} + \frac{f_9}{A_9 + C_9} \frac{A_{98} + C_{98}}{A_8 + C_8} A_{86}, \quad (2)$$

The subscripts ν and ν' ($\nu' < \nu$) are the vibrational states of OH before and after the corresponding process. The terms f_ν are the nascent distributions and describe the production efficiency of OH(ν) via the reaction $H+O_3$. Total radiative loss due to spontaneous emissions is considered by the Einstein coefficients A_ν (s^{-1}) which are the inverse radiative life times of OH(ν). The total loss rate C_ν (s^{-1}) is the sum of loss due to collisions with the air compounds (N_2 , O_2 , $O(^3P)$), including chemical reactions and physical quenching. The terms $A_{\nu\nu'}$ and $C_{\nu\nu'}$ represent the specific state-to-state transitions.

In the second step, chemical equilibrium of O_3 during night is assumed as follows:

$$k_1 [H][O_3] + k_2 [O(^3P)][O_3] = k_3 [O(^3P)][O_2][M], \quad (3)$$

meaning that O_3 loss due to H and $O(^3P)$ (left side) is balanced by O_3 formation via the three-body-reaction $O(^3P)+O_2+M$ (right side). Here, k_2 and k_3 are the corresponding rate constants of $O(^3P)+O_3$ and $O(^3P)+O_2+M$, respectively, while M being an air molecule and $[M]$ being the total number density of the air.

Finally, rewriting Eq. (1) enables the derivation of $[H]$ while $[O(^3P)]$ is calculated by substituting Eq. (3)

in Eq. (1) and rewriting the resulting term as follows:

$$[H] = \frac{OH(9-7) + OH(8-6) \text{ VER}}{G k_1 [O_3]}, \quad (4a)$$

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$$[O(^3P)] = \frac{OH(9-7) + OH(8-6) \text{ VER}}{G (k_3 [O_2][M] - k_2 [O_3])}. \quad (4b)$$

Air temperature and air pressure from SABER were used to calculate $[M]$, $[O_2]$ (VMR of 0.21), and

$[N_2]$ (VMR of 0.78) via the ideal gas law, and $[M]$ was then used to convert SABER O_3 VMR into $[O_3]$.

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The chemical reaction rates and physical quenching processes involved are described in Sect. 2.3. The values of $[O(^3P)]$ and $[H]$ were individually derived for each altitude. Finally, the obtained vertical profiles of $[O(^3P)]$ and $[H]$ were used to initialize the OH airglow model (see Sect. 2.3).

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It is apparent from Eq. (4a-b) that any changes applied to the input parameters (G , O_2 , O_3 , M , k_1 , k_2 , k_3) are balanced by the derived values of $[O(^3P)]$ and $[H]$, without assuming any a priori information of $[O(^3P)]$ and $[H]$. In contrast, $OH(9-7) + OH(8-6) \text{ VER}$ is not affected by the input parameters and therefore identical in every model run. However, the goal of this paper is to develop a model which does not only fit $OH(9-7) + OH(8-6) \text{ VER}$ observations but also reproduces the three other airglow measurements $OH(6-2) \text{ VER}$, $OH(5-3) + OH(4-2) \text{ VER}$, and $OH(3-1) \text{ VER}$. We have to further point out, that the relation between $[O(^3P)]$ and $OH(9-7) + OH(8-6) \text{ VER}$ is not linear since the function G also depends on $[O(^3P)]$, as represented by the terms C_v and C_{vv} . In fact, Eq. (4b) is a quadratic expression with respect to $[O(^3P)]$ but treated here as a linear one, making no substantial differences for small $[O(^3P)]$. Nevertheless, this issue is addressed in detail in Sect. 3.4.

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2.3 The OH airglow Base model

The model used in this study is based on the atmospheric chemistry box model Module Efficiently Calculating the Chemistry of the Atmosphere/Chemistry As A Box model Application (MECCA/CAABA-3.72f; Sander et al., 2011). The box model calculates the temporal evolution of chemical species inside a single air parcel of a certain pressure and temperature, making the model well suited for sensitivity studies. The CAABA/MECCA standard model was extended by several chemical reactions and physical quenching processes involving $OH(v)$ which are described in this section. The

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model was run until it reaches steady-state, defined by the agreement between the measured and modelled OH(9-7)+OH(8-6) VER.

240 The OH airglow model described in this section is referred to as “Base model” because it is the starting point of our model studies. But we have to point out that there is no such a thing as a commonly accepted OH airglow base model in the literature. The Base model takes into account all major formation and loss processes of OH(v) (Table 1) which are commonly used in other models in the literature and are assumed not to be seriously in error. The model comprises the production of OH(v)
245 via the chemical reaction H+O₃ as well as the deactivation due to spontaneous emission and the removal physical quenching and chemical reactions with N₂, O₂, and O(³P).

The chemical reactions H+O₃, O(³P)+O₃, and O(³P)+O₂+M were already included in the CAABA/MECCA standard model and their corresponding rates were taken from the latest Jet Propulsion Laboratory (JPL) report 18 (Burkholder et al., 2015). The reaction H+O₃ can populate
250 OH(v) at all vibrational levels $v \leq 9$ and the nascent distribution of OH(v) was taken from Adler-Golden (1997). The spontaneous emissions are given by the Einstein coefficients at 200 K (Xu et al., 2012). Deactivation of OH(v) by N₂ is assumed to occur via single-quantum quenching. The rates at room temperature for OH(v≤8) and for OH(v=9) were taken from Adler-Golden (1997) and Kalogerakis et al. (2011), respectively.

255 Quenching of OH(v) by O₂ is based on the values reported by Adler-Golden (1997, their Table 3) which comprise a combination of multi-quantum and single-quantum quenching. However, Adler-Golden (1997) applied a factor of ~1.5 to account for mesopause temperature based on comparisons between laboratory measurements at room temperature of OH(v=8)+O₂ and the corresponding rate inferred from OH(8-3) rocket observations in the mesopause region. But later experiments reported by Lacouisiere et al. (2003) and calculations by Caridade et al. (2002) suggest smaller values. The latter study further
260 indicates that the temperature dependence decreases for lower vibrational levels and becomes negligible for OH(v≤4). Consequently, the rates presented in Adler-Golden (1997) were scaled to room temperature measurements (v=1-6 Dodd et al., 1991; v=7 Knutsen et al., 1996; v=8 Dyer et al., 1997; v=9 Kalogerakis et al., 2011), and afterwards a factor of 1.1 for OH(v≥6) and 1.05 for OH(v=5) was

265 applied.

The removal of OH(v) via collisions with O(³P) is included by using a combination of multi-quantum quenching (Caridade et al., 2013, their Table 1) and chemical reactions (Varandas, 2004). The rates were obtained from quasi-classical trajectory calculations at 210 K, approximately matching mesopause temperature.

270 As described in the previous section, the OH airglow model is adjusted to fit OH(9-7)+OH(8-6) VER, OH(6-2) VER, OH(5-3)+OH(4-2) VER, and OH(3-1) VER. Thus, the model cannot provide information about OH(v≤2). It further treats OH(v=9) and OH(v=8) as well as OH(v=5) and OH(v=4) as a single level and the corresponding deactivation channels presented in Table 2 and 3 should be viewed more critically.

275 **3 Results and discussions**

Figure 1 displays vertical profiles of a) OH(6-2) VER, b) OH(5-3)+OH(4-2) VER, and c) OH(3-1) VER, comparing the observations (black squares) with the corresponding Base model output (red line). The model results of OH(6-2) VER and OH(3-1) VER are a 4 km running average to take the averaging kernels of SCIAMACHY measurements into account. The Base model approximately matches the general shape of the measured profiles but overestimates the three OH airglow measurements at the altitude of maximum VER. A closer look at the relative differences shows that the ratio model/observation at the altitude of maximum VER is about 2.0, 1.2, and 1.3 for OH(6-2), OH(5-3)+OH(4-2), and OH(3-1), respectively. Furthermore, these ratios increase with decreasing altitude, indicating that the overestimation of the Base model might be associated with O₂ quenching.

280
285 The differences between Base model and observations are quite substantial in case of OH(6-2) VER. This implies a general problem of the rates or schemes included in the Base model, requiring a detailed error analysis. The focus was set on potential error sources of OH(6-2) VER because the relative differences between model and measurements are largest compared to the other two OH transitions, and secondly because changes of OH(v=6) will affect the lower vibrational levels, but not vice versa.

290 **3.1 Potential error sources of OH(6-2) VER in the Base model**

Based on the results presented in Fig. 1, the potential error source has to have an effect on the entire height interval and must have a stronger impact on OH(6-2) compared to the other two OH transitions. We further focus on quantities with large uncertainties. For the latter reason, temperature is excluded as possible source because to account for a reduction of OH(6-2) VER by a factor of 2, temperature must
295 be increased by more than 20 K (not shown here). Such a large error is very unlikely considering that a zonal mean climatology (2003-2011) is used here.

Since the overestimation of the Base model is especially large for OH(6-2) VER, an impact of the Einstein coefficient of the corresponding transition must be considered. Regarding this aspect, we have to point out that studies based on HITRAN 2004 data set should be viewed more critically, because of
300 erroneous OH transition probabilities. The Einstein coefficients used in this study were recently recalculated (Xu et al., 2012, their Table A1) and correspond to a temperature of 200 K, which is very close to mesopause temperature. Furthermore, these Einstein coefficients are consistent with the values of the HITRAN 2008 data set (Rothman et al., 2009). However, there are several other data sets of Einstein coefficients found in literature that might lead to different results. We therefore carried out sensitivity
305 runs, using the Einstein coefficients reported by Turnbull and Lowe (1989), Nelson et al. (1990), van der Loo and Groenenboom (2007), Xu et al. (2012; =Base model), and Brooke et al. (2016). The corresponding results are presented in Figure 2 and show considerably large differences in case of OH(6-2) VER which are about a factor of 4 between the highest and lowest model output. In contrast, the individual simulations of OH(5-3)+OH(4-2) VER and OH(3-1) VER are rather consistent and vary
310 only by ~10 %. These results emphasize that the choice of the Einstein coefficients is a potential error source for higher quanta transitions.

Regarding the credibility of the Einstein coefficients, it is generally assumed that the calculation improve with time. However, this is not necessarily true at quanta changes >2 because it all depends on how good the representation of the Hamiltonian for the OH molecule is, that is used to solve the
315 Schrödinger equation. Multi quanta transitions >2 quanta have small Einstein coefficients and are generally hard to model and calculate. The assessment of the Einstein coefficients requires a detailed analysis of the corresponding calculations, which is beyond the scope of this study. We therefore cannot exclude the values used in the Base model as a potential error source, but we also think that our choice

of the Einstein coefficients from Xu et al. (2012) is reasonable. Additionally, these values represent approximately the average model output of all five data sets considered here, while the model results based on Nelson et al. (1990) and van der Loo and Groenenboom (2007) represent the variability. Thus, we will not replace the Einstein coefficients by Xu et al. (2012) in our model but keep in mind that they might be too large.

Furthermore, the best agreement between the observations and the model was obtained by applying the Einstein coefficients reported by van der Loo and Groenenboom (2007). But even in this case, the model still overestimates the observations of all OH transitions in the altitude region between ~80 km and ~86 km. This pattern strongly supports the suggestion stated above that the rates and schemes associated with OH(v)+O₂ are incorrect.

The nascent distribution of the excited OH states of the chemical reaction H+O₃ was observed in several studies and all of them agree that OH(v) is primarily formed in the vibrational levels v=8 and v=9 (e.g. Charters et al., 1971; Streit and Johnston, 1976; Ohoyama et al., 1985; Klenerman and Smith, 1987). The values used in the Base model were taken from Adler-Golden (1997) which are based on measurements reported by Charters et al. (1971) and agree with values obtained by Klenerman and Smith (1987) and Streit and Johnston (1976). The values found by Ohoyama et al. (1985) show some differences, but according to Klenerman and Smith (1987), their results are fundamentally flawed. This also affects the nascent distribution used by Mlynchak and Solomon (1993) which is an average of Charters et al. (1971), Ohoyama et al. (1985), and Klenerman and Smith (1987).

Therefore, we think that our nascent distribution used here is likely not a serious error source. However, minor errors might be introduced by extrapolating the nascent distribution to lower vibrational levels as it was done for the values used in our study (Adler-Golden, 1997). It is also possible that part of the nascent value of OH(v=6) is not due to direct production via H+O₃ but results from contributions of OH(v≥7). In order to test the potential impact of the OH(v=6) nascent value on OH(6-2) VER, we assumed an extreme scenario by reducing the OH(v=6) nascent value from 0.03 to zero. But the corresponding results of OH(6-2) VER of the Base model run (not shown here) are only about 15 % lower compared to the values presented in Fig. 1. Further sensitivity runs also showed that an increase of the ratio f₉/f₈ is associated with a decrease of modelled OH(6-2) VER but even the extreme case of

$f_9=1$ and $f_8=0$ could not account for a factor of 2. Note that changes of the overall rate constant of $\text{H}+\text{O}_3$ affect all considered OH transitions in a similar way, Thus, we conclude that direct production of OH(v) is unlikely to be the reason for the overestimation of OH(6-2) VER by the Base model.

350 The physical removal of OH(v) by N_2 is included as single-quantum relaxation which is supported by theoretical studies (Shalashilin et al., 1992; Adler-Golden, 1997). Assuming a sudden death scheme with the same overall deactivation rates resulted in a decrease of simulated OH(6-2) VER by less than 10 % at the altitude of maximum VER. The total deactivation rate for OH(v=9) used here is about 1.5 times higher than the one suggested by Adler-Golden (1997) but the difference between the corresponding
355 model OH(6-2) VERs is negligible (<1 %). There are two studies reporting temperature dependence of N_2 quenching (Shalashilin et al., 1992; Burt and Sharma, 2008), both agreeing with measurements at room temperature. However, the calculations of the former study imply slower quenching rates at mesopause temperature compared to their respective values at room temperature whereas the latter publication indicates the opposite behaviour, reporting a ratio between the rate at 200 K and 300 K of
360 approximately 1.7 for OH(v=8) and 1.3 for OH(v=9). These factors are generally supported by López-Puertas et al. (2004) which applied an empirically determined factor of 1.4 to the rates of Adler-Golden (1997) to account for mesopause temperature. Since the temperature dependence is still uncertain, we tested both possibilities. We increased and decreased the overall OH(v)+ N_2 quenching rates by a factor of 1.5 which led to higher or lower OH(6-2) VERs by about 5 %. Therefore, N_2 is too inefficient as a
365 OH(v) quenching partner to cause differences of OH(6-2) VER of a factor of 2.

The overall rate and exact pathways of OH(v)+ $\text{O}(^3\text{P})$ are also still not known well enough but $\text{O}(^3\text{P})$ has nearly no influence on OH(v) at altitudes below 85 km. It therefore cannot be the only reason for the differences presented in Fig. 1. Consequently, deactivation by O_2 is the only remaining candidate which has a crucial influence on OH(v) throughout the entire height interval. Therefore, we will first focus on
370 OH(v)+ O_2 (Sect. 3.2) before investigating a potential influence of $\text{O}(^3\text{P})$ on OH(v) in Sect. 3.3.

3.2 Deactivation of OH(v) by O_2

The overestimation of OH(6-2) VER by the Base model can be generally corrected either by slower rates of OH(v=9,8,7)+ O_2 or by a faster rate of OH(v=6)+ O_2 . The overall deactivation of OH(v=9) was

375 measured by Chalamala and Copeland (1993) and they recommended a value of $2.1 \times 10^{-11} \text{ cm}^3 \text{ s}^{-1}$. This result was later confirmed by Kalogerakis et al. (2011), reporting a rate of $2.2 \times 10^{-11} \text{ cm}^3 \text{ s}^{-1}$. The rates for $\text{OH}(v=8,7,6)+\text{O}_2$ are each based on a single study only ($v=8$ Dyer et al., 1997; $v=7$ Knutsen et al., 1996; $v=6$ Dodd et al., 1991). But at least to our knowledge, there are no signs that the rates of $\text{OH}(v=9,8,7,6)+\text{O}_2$ are fundamentally flawed. In order to test the impact of the individual rates on $\text{OH}(6-2)$ VER, we carried out sensitivity runs by varying the overall rates within their recommended 2σ errors. Thus, we reduced the values of $\text{OH}(v=9,8,7)+\text{O}_2$ to $16 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, $7 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, and $5 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, respectively, while the rate of $\text{OH}(v=6)+\text{O}_2$ was increased to $4.5 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$. But even under this favoured condition, the Base model output of $\text{OH}(6-2)$ VER decreased only by a factor of 1.5, still not close to the required difference of a factor of 2. Additionally, the assumed scenario is rather unlikely since the overall rates were obtained by independent studies.

385 The possibility of a systematic offset of $\text{OH}(v \leq 6)+\text{O}_2$ rates, which are based on the single study (Dodd et al., 1991), is also excluded because of the very good agreement of this $\text{OH}(v=2)+\text{O}_2$ rate with the value obtained by Rensberger et al. (1989). Furthermore, when we increased the $\text{OH}(v \leq 6)+\text{O}_2$ rates by a factor of 3, the Base model approximately fits $\text{OH}(6-2)$ VER and $\text{OH}(3-1)$ VER but underestimates $\text{OH}(5-3)+\text{OH}(4-2)$ VER by more than 30 %. Temperature dependence also affects the O_2 deactivation rates used here. But the factor to account for mesopause region temperature is suggested to be lower than 1.3 (Lacouisiere et al., 2003; Cadidade et al., 2002) which has a weaker impact on $\text{OH}(6-2)$ VER than the scenarios considered above.

Consequently, when applying the standard deactivations rates and schemes found in the literature, neither errors of the overall rates nor uncertainties of the temperature dependence can give a reasonable explanation of the overestimation of $\text{OH}(6-2)$ VER Base model output shown in Fig. 1a. Since the overall rates were actually measured while the deactivation schemes are solely based on theoretical considerations, it is more convincing that the potential error source lies within $\text{OH}(v)+\text{O}_2$ deactivation scheme rather than in the deactivation rates.

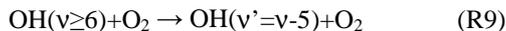
395 In order to considerably reduce $\text{OH}(6-2)$ VER, we assumed an extreme scenario and substituted the multi-quantum relaxation ($\text{OH}(v)+\text{O}_2 \rightarrow \text{OH}(v' < v)+\text{O}_2$) in the Base model by a sudden death

(OH(v)+O₂→OH+O₂) approach. This new model is referred to as “O₂ SD model” and the corresponding results are displayed in Fig. 3 as red lines, showing that the simulated OH(6-2) VER matches the observations within the error bars below 85 km and above ~92 km. The model still overestimates the measurements in the altitude region ~90 km, which might be related to O(³P) quenching (see Sect. 3.3). The O₂ SD model output for the other two OH transitions (Fig. 3b-c) is clearly too low, implying that OH(v)+O₂ quenching cannot occur via sudden death alone. We also conclude that the contribution of higher excited states OH(v≥7) to OH(v=6) must be negligible or even zero and these higher states are suggested to primarily populate lower vibrational levels OH(v≤5). Therefore, OH(v)+O₂ has to occur via multi-quantum quenching because in case of single-quantum deactivation the contribution of OH(v≥7) to OH(v=6) is considerably larger than zero.

According to Finlayson-Pitts and Kleindienst (1981), OH(v) might be relaxing to v'=v-5 while the excess energy is transferred to form O₂(b¹Σ). This vibration-to-electronic energy transfer was also mentioned by Anlauf et al. (1968) and is supported by the close energy match of the transition from OH(v=9) to OH(v=4) and from O₂(X³Σ) to O₂(b¹Σ) of about 36.6 kcal mol⁻¹ and 37.5 kcal mol⁻¹, respectively. Although there is no experimental support of this deactivation pathway, this approach gives a reasonable explanation for the observed pattern in our study and OH(v) as a potential source of excited O₂, as discussed in Howell et al. (1990) and Murtagh et al. (1990). However, evaluating whether the product is really O₂(b¹Σ) or another excited O₂ state is beyond the scope of this study. Thus, we concluded that deactivation of OH(v) by O₂ has to satisfy the following condition:

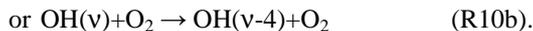


while we further assume that the pathway



is the preferred deactivation channel.

In order to test whether R9 could be the only pathway of R8 we assumed multi-quantum relaxation via:



If R10a is integrated in the model (Fig. 3b-c, O₂ v-5 model), the corresponding model output at altitudes

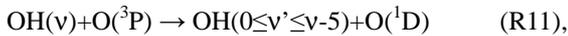
<90 km is only about 10 % below the observations of OH(5-3)+OH(4-2) VER and approximately matches OH(3-1) VER measurements within the error bars. The underestimation of the OH(5-3)+OH(4-2) VER measurements by the model could be attributed to minor errors of the OH(v)+O₂ overall rates in combination with a slightly different OH(v) branching of H+O₃. Therefore, we cannot completely rule out R10a as a possible solution, even if there are still some differences between the modelled and the observed OH VER. Replacing R10a by R10b in the model (Fig. 3b-c, O₂ v-4 model) results in an overestimation of the observations of OH(5-3)+OH(4-2) VER and OH(3-1) VER by about 20 % to 30 %, and consequently this assumption is not further considered as a potential solution.

The results shown in Fig. 3 suggest that the OH airglow model is not able to reproduce the three OH airglow observations when sudden death or simplified multi-quantum schemes for OH(v)+O₂ are applied. But the O₂ v-5 model output is quite close to the measurements, suggesting that R9 might be the dominating deactivation channel within a multi-quantum relaxation scheme in accordance with R8. We therefore included these two conditions in the so-called “O₂ best fit model” and the results are displayed in Fig. 4. The corresponding branching ratios for the individual pathways are summarized in Table 2.

The simulated OH airglow fits well with the three OH airglow observations within the error bars below 85 km. In the altitude region above 85 km, it is seen that the model still overestimates OH(6-2) VER while OH(3-1) VER is indicated to be slightly underestimated. Furthermore, this pattern is not seen in OH(5-3)+OH(4-2) VER and therefore could be attributed to deviations due to the different satellite/instrument configurations between TIMED/SABER and ENVISAT/SCIAMACHY. But since this behaviour only occurs in the upper part of the vertical profiles and is not seen throughout the entire height interval, it is more likely related to O(³P) quenching.

3.3 Deactivation of OH(v) by O(³P)

Only recently, Sharma et al. (2015) proposed a new pathway of OH(v)+O(³P) by providing a direct link between higher and lower vibrational levels via:



with the vibrationally independent reaction constant $k_{11} = 2.3 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$. While the value of $k_{11}(v=9)$

455 is based on measurements (Kalogerakis et al., 2011; Thiebaud et al., 2010) and on calculations
(Varandas, 2004), the values for $k_{11}(v=5, 6, 7, 8)$ are only assumed to be identical to $k_{11}(v=9)$ and should
be viewed more critically.

We adapted R11 in the “O₂ best fit O(³P) v-5 model” in such a way that the product is OH(v'=v-
5)+O(¹D) and the results obtained are displayed as blue lines in Fig. 5. Comparisons for OH(6-2) VER
460 in Fig. 5a show an underestimation of the model at altitudes >85 km. A sensitivity study was carried out
that showed that the impact of OH(v=9,8,7)+O(³P) on OH(6-2) VER is negligible. This is reasonable
because these three upper states only indirectly influence OH(6-2) via R11. Consequently, our analysis
suggests a lower value of $k_{11}(v=6)$ and best agreement between model output and OH(6-2) VER
observations was obtained for an overall rate of approximately $0.8 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$.

465 In case of OH(5-3)+OH(4-2) VER, presented in Fig. 5b, the new approach leads to a weak
underestimation of the observations by the model in the altitude region above 85 km, even if
OH(v=9)+O(³P) of R11 solely populates OH(v=4). The model results are most sensitive to $k_{11}(v=5)$, and
therefore this rate might be too fast. Considering our best fit value obtained for $k_{11}(v=6)$, it is indicated
that $k_{11}(v)$ decreases with decreasing vibrational level and this feature is discussed below in more detail.
470 Thus, an upper limit of $k_{11}(v=5) < k_{11}(v=6)$ is recommended and the actual rate coefficient has to balance
the direct contribution of OH(v=9) to OH(v=4) via R11. Investigating another scenario of $k_{11}(v=5)$
being zero showed that the branching of OH(v=9) to OH(v=4) has to be at least about 0.6 which
corresponds to a rate of a $\sim 1.4 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$.

The assumption that $k_{11}(v)$ decreases at lower vibrational levels is supported by the overall rate of
475 OH(v=7)+O(³P)→OH(v')+O(¹D) at mesopause temperature which is suggested to be on the order of
 $0.9-1.6 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$ (Thiebaud et al., 2010; Varandas, 2004). At least to our knowledge, the total rate
of OH(v=8)+O(³P)→OH(v')+O(¹D) was not measured. Nevertheless, results reported by Mlynczak et
al. (2018) and Panka et al. (2017, 2018) indicate that this rate might be slower than the value of 2.3×10^{-10}
 $\text{ cm}^3 \text{ s}^{-1}$ suggested by Sharma et al. (2015). This is also in agreement with our findings here, because
480 applying $2.3 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$ for $k_{11}(v=9,8)$ results in non-physical [O(³P)] values above 90 km. The
corresponding value of [O(³P)] e.g. at 95 km is about 1.25 times larger than SABER [O(³P)] 2013

(Mlynczak et al., 2013a) which in turn is about 1.15 times larger than the upper limit of $[O(^3P)]$ (Mlynczak et al., 2013b, their Fig. 4). This results in a factor of $1.15 \times 1.25 = 1.44$ (=44 %) above the upper limit and cannot be explained by the uncertainty of the $[O(^3P)]$ profile derived here (40 %, see
485 Sect. 3.4). In order to obtain reasonable $[O(^3P)]$ values, it was necessary to lower the rate of $k_{11}(v=8)$ to $1.8 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$, and we therefore recommend $k_{11}(v=8) \leq 1.8 \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$ as an upper limit to derive physically allowed $[O(^3P)]$ values.

It is seen in Fig. 5c that observations and O_2 best fit $O(^3P)$ v-5 model output of OH(3-1) VER are in agreement within the corresponding measurement errors but the model values seem to be slightly too
490 low at heights >85 km. In this altitude region, simulated OH(3-1) VER is most influenced by $OH(v=9,8)+O(^3P)$ of R11 because both vibrational levels can directly populate $OH(v=3)$. However, not much is known about the individual branching ratios of R11 except that $OH(v=9)+O(^3P) \rightarrow OH(v=3)+O(^1D)$ is an important deactivation channel but not necessarily the dominating one (Kalogerakis et al., 2016). These authors suggested a rate of $2.3(\pm 1.0) \times 10^{-10} \text{ cm}^3 \text{ s}^{-1}$ and
495 noted that this rate might be slower due to the involvement of excited surfaces. This generally agrees with our results presented here because the O_2 best fit $O(^3P)$ v-5 model only considers a contribution of $OH(v=8)$ to $OH(v=3)$ and the underestimation indicated in Fig. 5c could be attributed to the missing channel $OH(v=9)+O(^3P) \rightarrow OH(v=3)+O(^1D)$. The conclusions drawn from comparisons between three different airglow observations and our model studies with respect to $OH(v)+O(^3P)$ quenching are
500 summarized in Table 3.

Finally, all these findings presented in Table 2 and 3 were adapted in the “Best fit model” (Fig. 5, red lines), resulting in an overall agreement between model output and measurements within the corresponding errors. Note that $k_{11}(v=7)$ used here is the average of the lower and upper limits derived from Thiebaud et al. (2010) and Varandas (2004) which is unlikely to be seriously in error. Furthermore,
505 we have to point out that lowering $k_{11}(v=8)$ does only impact the $[O(^3P)]$ and $[H]$ derived here but does not affect the general conclusions drawn in this section.

The empirically determined solution presented here implies that the contribution of $OH(v=9)$ to $OH(v=8)$ via quenching with $O(^3P)$ is close to zero (see Table 1 and this section). In contrast, the model

described in Mlynczak et al. (2018) assumes single-quantum relaxation
510 $(\text{OH}(v=9)+\text{O}(\text{}^3\text{P})\rightarrow\text{OH}(v=8)+\text{O}(\text{}^3\text{P}))$ to get the global annual energy budget into near balance. But
applying this approach in our OH model (same total rate of $3\times 10^{-10}\text{ cm}^3\text{ s}^{-1}$ and varying the rates for
 $\text{OH}(v\leq 8)+\text{O}(\text{}^3\text{P})$) leads to a considerable overestimation of OH(6-2) VER. Additionally, the shape of
simulated OH(5-3)+OH(4-2) VER slightly mismatches the observed OH(5-3)+OH(4-2) VER above 90
km (not shown here). Based on these sensitivity runs, we conclude that at least part of OH(v=9)+O(³P)
515 channel has to be deactivated via multi-quantum quenching. This is supported by the results presented
by Panka et al. (2017) which adjusted an OH airglow model to fit night-time CO₂(v₃) emissions at 4.3
μm. However, this study reported empirically determined rates for OH(5≤v≤8)+O(³P) generally higher
than the rates obtained in this work. But these differences might be attributed to their faster values of
OH(v)+O₂ because they seem to have falsely assumed that the rates of Adler-Golden (1997) do not take
520 mesopause temperature into account. Thus, we think that their rates of OH(v)+O₂ are too high, at least
by a factor of ~1.5. Since they performed an empirical study, it is not possible to estimate how much
this issue affects the rates of OH(5≤v≤8)+O(³P). But we know from our work that higher rates of
OH(v)+O₂ lead to higher values of OH(6-2) VER, OH(5-2)+OH(4-2) VER, and OH(3-1) VER which
can be generally balanced by higher rates of OH(5≤v≤8)+O(³P). Considering our comparisons with
525 these two studies, we think that the rates of OH(v)+O(³P) should be investigated in more detail in future
studies as this rate has a huge impact on derived values of [O(³P)] (Panka et al., 2018).

3.4 Derived profiles of [O(³P)] and [H]

Figure 6 displays the vertical profiles of [O(³P)] and [H] obtained by the Best fit model in comparison
with the results derived from SABER OH(9-7)+OH(8-6) VER only (Mlynczak et al., 2018). The
530 [O(³P)] profiles seen in Fig. 6a agree below 85 km but the Best fit model shows gradually larger values
in the altitude region above. Between 85 km and 95 km, these larger values are caused by the different
deactivation rates and schemes of OH(v)+O(³P), agreeing with general pattern reported in Panka et al.
(2018). We have to point out that other studies (e.g. von Savigny and Lednyts'ky, 2013) observed a
pronounced [O(³P)] maximum of about $8\times 10^{11}\text{ cm}^{-3}$ at 95 km. The [O(³P)] derived here indeed shows

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535 similar values at 95 km but a maximum is not seen. Nevertheless, the $[O(^3P)]$ in our study obtained above 95 km looks rather unexpected and possible reasons are discussed below.

The night-time $[H]$ derived in this study shows similar pattern as SABER $[H]$, including the maximum at 80 km. But Best fit model $[H]$ is systematically larger than SABER $[H]$ by a factor of approximately 1.5. This is primarily caused by our faster $OH(v=8)+O_2$ rate compared to the rate applied in Mlynckzak
540 et al. (2018). Similar to the comparisons with $[O(^3P)]$, Best fit model $[H]$ results also shows unexpected patterns above 95 km.

The quality of the derived profiles is primarily affected by three different uncertainty sources. The first source includes uncertainties due to the rates of chemical and physical processes as well as the background atmosphere considered in the Best fit model. We assessed the 1σ uncertainty by assuming
545 uncorrelated input parameters. Adler-Golden (1997) did not state any uncertainties for f_9 and f_8 but these values should be similar to the uncertainty of f_8 derived by Klenerman and Smith (1987). Therefore, we applied an uncertainty of 0.03 for f_9 and f_8 . In case of the Einstein coefficient, we adapted an uncertainty of 30 %, which is based on the five sets of Einstein coefficients analyzed in Sect. 3.1. Note that larger uncertainties only occur for multi quanta transitions >2 quanta. But $[O(^3P)]$ and $[H]$ were calculated
550 from the transition $OH(9-7)+OH(8-6)$ where the agreement is better. All the other 1σ uncertainties of the input parameters were taken from their respective studies.

Recent comparisons between MIPAS O_3 and SABER O_3 derived at $9.6 \mu m$ were performed by Lopez-Puertas et al. (2018). The authors showed that night-time O_3 from SABER is slightly larger than night-time O_3 obtained from MIPAS in the altitude region 80-100 km over the equator (their Fig. 8 and 10)
555 but these differences are within the corresponding errors. Thus, at least to our knowledge there is no conclusive evidence stating that SABER night-time O_3 is generally too large. Nevertheless, we considered an uncertainty of O_3 of about 10 % (Smith et al., 2013). The uncertainty of SABER temperature was estimated to be lower than 3 % (Garcia-Comas et al., 2008) while the total uncertainty of SABER $OH(9-7)+OH(8-6)$ VER was assumed to be about 6 % (see Sect. 2.1.2). The total 1σ
560 uncertainty was obtained by calculating the root-sum-square of all individual uncertainties. The results of 1σ uncertainty of $[O(^3P)]$ and $[H]$ derived by the Best fit model are shown as error bars in Fig. 6. The

error bars of SABER [O(³P)] and [H] were adapted from the corresponding publication.

In case of the Best fit model [O(³P)] profile, the 1 σ uncertainty varies between 30 % and 40 %, depending on altitude. The individual contributions of the input parameters to the total 1 σ uncertainty are considerably different. Einstein coefficients and nascent distribution each account for about 10 % and 5 %, respectively, throughout the entire height interval. The influence of the collision rates is about 5 % and gradually decreases to zero with increasing altitude. In contrast, the chemical reaction rates k_2 and k_3 account for ~80 % to ~85 % of the overall 1 σ uncertainty of the derived [O(³P)] profiles. The total 1 σ uncertainty of [H] varies between 25 % and 40 % with k_1 being the major uncertainty source (~80 %) below 85 km. In higher altitude regions, the impact due to uncertainty of [O(³P)] becomes gradually more important and both k_1 and [O(³P)] each contribute close to one half to the overall uncertainty at altitudes >95 km. We further assumed a worst case scenario (not shown here), meaning that all uncertainties of the input parameters contribute to either higher or lower [O(³P)] values, obtaining a worst case 1 σ uncertainty of approximately 80 % for [O(³P)] and about 65 % for [H]. However, it is more likely that the uncertainties are uncorrelated since they originate from independent measurements.

The second aspect influencing the quality of the derived profiles is the assumption of chemical equilibrium of O₃, represented by Eq. (3). This issue was recently investigated by Kulikov et al. (2018), which carried out simulations with a 3-D chemical transport model and demonstrated that a wrongly assumed chemical equilibrium of O₃ may lead to considerable errors of derived [O(³P)] and [H]. In order to test the validity of chemical equilibrium of O₃ locally, the authors suggested that OH(9-7)+OH(8-6) VER has to exceed 10×G×B, with B including several chemical reaction rates involving O_x and HO_x species. Note that this criterion requires simultaneously performed temperature and OH airglow measurements. Furthermore, this criterion is based on the assumption that the impact of atmospheric transport on chemical equilibrium of O₃ is negligible. Since our experiments fit these conditions, we applied their suggested limit and found that in our case chemical equilibrium of O₃ is valid above 80 km. We have to point out that the term “chemical equilibrium of O₃” refers to O₃ that does not deviate more than 10 % from O₃ in chemical equilibrium (Kulikov et al., 2018, their Eq. 2).

Assuming that O₃ is always 10 % ~~greater or lesser than~~ O₃ in chemical equilibrium introduces an

Gelöscht: above

Gelöscht: below

590 uncertainty of about 10 % at 80 km and 20 % at 95 km, additionally to the total uncertainty of $[O(^3P)]$
and $[H]$ estimated above. However, such a worst case scenario is rather unlikely while it is more
realistic that O_3 actually varies around its chemical equilibrium concentration. Thus, an over- and
underestimation of derived $[O(^3P)]$ and $[H]$ are assumed to compensate each other. Consequently, we
conclude that the impact on the total uncertainty of $[O(^3P)]$ and $[H]$ due to deviations from chemical
595 equilibrium of O_3 is negligible, but only because the previously used criterion $(OH(9-7)+OH(8-6))$
 $VER > 10 \times G \times B$ is valid.

The last problem lies in the fact that the approach used here (see Sect. 2.2) has to be applied to
individual OH airglow profiles to derive $[O(^3P)]$ and $[H]$ correctly. However, the individual scans of
OH(6-2) were too noisy to analyze single profiles and we therefore used climatology for all input
600 parameters. By investigating individual OH airglow profiles, we would derive individual $[O(^3P)]$
profiles and eventually average them to the mean $[O(^3P)]$ profile. While in our case, we directly derive
the mean $[O(^3P)]$ profile. This makes no difference as long as the relation between OH airglow and
 $[O(^3P)]$ is a linear one. But Eq. (4b) shows that the relation between $[O(^3P)]$ and $(OH(9-7)+OH(8-6))$
VER is only approximately linear because G also depends on $[O(^3P)]$, as represented by the terms C_v
605 and C_{vv} . The linearity between $(OH(9-7)+OH(8-6))$ VER and $[O(^3P)]$ of an air parcel with a certain
temperature and pressure is solely controlled by $[O(^3P)] \times G$. Note that $[H]$ too is affected by this non-
linearity issue since $[H]$ depends on G (Eq. (4a)). Thus, derived $[H]$ values are only reliable as long as
the derived $[O(^3P)]$, and as a consequence G , is not seriously in error.

In order to test the linearity, $[O(^3P)] \times G$ was plotted as a function of $[O(^3P)]$ and the corresponding
610 results for Best fit model at five different heights are presented in Fig. 7. It is seen that the relation
between $[O(^3P)]$ and $[O(^3P)] \times G$ or $(OH(9-7)+OH(8-6))$ VER, respectively, is linear for small values of
 $[O(^3P)]$, while a non-linear behaviour becomes more pronounced for larger values of $[O(^3P)]$.
Furthermore, the starting point of the behaviour is shifted to lower $[O(^3P)]$ values at higher altitudes. In
order to estimate this threshold, we performed a visual analysis and determined an upper limit of
615 $[O(^3P)]$ before non-linearity of $[O(^3P)] \times G$ takes over. The approximated upper limits are added as
dashed lines in Fig. 7. Finally, an $[O(^3P)]$ value at a certain altitude is assumed to be true if this value is
below the corresponding upper limit of $[O(^3P)]$. Otherwise, it should be viewed more critically. This

was done for each altitude and we found that the $[O(^3P)]$ and $[H]$ profiles presented in Fig. 6 are plausible in the altitude region <95 km. In combination with the estimation of chemical equilibrium of O_3 and the maximum of physically allowed $[O(^3P)]$, we think that the $[O(^3P)]$ and $[H]$ derived by the Best fit model are reasonable results between 80 km and 95 km. Note that these altitude limits do not affect the results with respect to $OH(v)+O_2$ and $OH(v)+O(^3P)$ presented in the Sect. 3.2 and 3.3.

4 Conclusions

We presented a zero dimensional box model which fits the VER of four different OH airglow observations, namely TIMED/SABER $OH(9-7)+OH(8-6)$ and $OH(5-3)+OH(4-2)$ as well as ENVISAT/SCIAMACHY $OH(6-2)$ and $OH(3-1)$. Based on a night-time mean zonal mean climatology of co-location measurements between 2003 and 2011 at $0^\circ-10^\circ$ N, we found that I) $OH(v)+O_2$ is likely to occur via multi-quantum deactivation while $OH(v \geq 7)$ primarily contribute to $OH(v \leq 5)$ and might prefer deactivation to $OH(v'=v-5)+O_2$. This relaxation scheme generally agrees with results reported in Russell et al. (2005) but is considerably different to the commonly used scheme suggested by Adler-Golden (1997). We further found II) general support for the new pathway $OH(v)+O(^3P) \rightarrow OH(v')+O(^1D)$ proposed by Sharma et al. (2015) but suggest slower total loss rates of $OH(v=8,7,6,5)+O(^3P)$. Additionally, hints for a favoured deactivation to $OH(v'=v-5)+O(^1D)$ are obtained.

We have to stress that we performed an empirical model study and the total rates and deactivation channels suggested here heavily depend on the OH transitions considered. Including additional OH transitions, like $OH(9-4)$, $OH(8-3)$, and $OH(5-1)$ from the Optical Spectrograph and InfraRed Imager System (OSIRIS) on board the Odin satellite, might result in other values and deactivation schemes. This could be a subject of a future study. Also note that the Einstein coefficients used here might be in error (see Sect.3.1; Fig. 2). This does not affect the two general conclusions drawn above but would impact the empirically derived rates.

Furthermore, our OH airglow model is based on the transitions $OH(9-7)+OH(8-6)$, $OH(6-2)$, $OH(5-3)+OH(4-2)$, and $OH(3-1)$ only. Therefore, our model does not provide any information of $OH(v \leq 2)$. It

645 further cannot distinguish between OH($v=5$) and OH($v=4$) as well as OH($v=9$) and OH($v=8$),
respectively, and errors in OH($v=5$) and OH($v=9$) might be compensated by errors in OH($v=4$) and
OH($v=8$) or vice versa. Consequently, the rates of the individual deactivation channels presented in
Table 2 and Table 3 should be viewed as a suggestion only. In particular, the rate of
OH($v=9$)+O(3 P) \rightarrow OH($v=3$)+O(1 D is about 3 times slower than the lower limit reported by Kalogerakis
et al. (2016). But these issues will only be solved eventually when future laboratory experiments
650 provide the corresponding OH(v)+O₂ and OH(v)+O(3 P) relaxation rates and deactivation channels.
Nevertheless, we have to emphasize that the shortcomings of our model do not affect the two main
conclusions drawn in this study.

Justified by a nearly linear relation between [O(3 P)] and OH(9-7)+OH(8-6) VER, the physically allowed
upper limit of [O(3 P)], and also considering the chemical equilibrium of O₃, we conclude that the
655 [O(3 P)] and [H] profiles derived by the Best fit model are plausible in the altitude range from 80 km to
95 km. The corresponding 1 σ uncertainty due to uncertainties of chemical reactions and physical
processes varies between 35 % and 40 % ([H]) and between 30 % and 40 % ([O(3 P)]), depending on
altitude.

The [H] derived here is systematically larger by a factor of 1.5 than SABER [H] reported in Mlynczak
660 et al. (2018) which is primarily attributed to their slower OH($v=8$)+O₂ rate. Our [O(3 P)] values in the
altitude region below ~87 km are in agreement within the corresponding errors with the results found in
Mlynczak et al. (2018) and Zhu and Kaufmann (2018) but are lower than the values presented in Panka
et al. (2018). However, we think that the results of the latter study are too large because the authors
falsely assumed too fast OH(v)+O₂ rates. In the altitude region above ~87 km, the [O(3 P)] shown here is
665 generally larger than the values reported in these three studies up to a factor 1.5 to 1.7. These
differences are attributed to the faster rates and different deactivation channels of OH(v)+O(3 P).
Therefore, it is indicated that we might overestimate [O(3 P)] above >87km and we suggest that our
results should be interpreted as an upper limit. However, a final conclusion cannot be drawn at this point
due the large uncertainties of the rates assumed to derive [O(3 P)].

670 **Data availability**

The data used in this study are open for public. The TIMED/SABER data can be downloaded from <http://saber.gats-inc.com/data.php> while ENVISAT/SCIAMACHY can be accessed by getting in contact with Christian von Savigny (csavigny@physik.uni-greifswald.de).

Author contribution

675 MS initialized and supervised the study. CVS retrieved the SCIAMACHY data. TF performed the model runs and wrote the final script with contributions from all co-authors.

Competing interests

The authors declare that they have no conflict of interest.

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References

- Adler-Golden, S.: Kinetic parameters for OH nightglow modeling consistent with recent laboratory
685 measurements, *J. Geophys. Res.*, 102, 19,969–19,976, doi:10.1029/97JA01622, 1997.
- Adreus, D. G., Holton, J. R., and Leovy, C. B. (Eds.): *Middle Atmosphere Dynamics*, Academic Press, Orlando, USA, 1987.
- Anlauf, K. G., MacDonald, R. G., and Polanyi, J. C.: Infrared chemiluminescence from H+O₃, at low pressure, *Chem. Phys. Lett.*, 1, 619-622, doi:10.1016/0009-2614(68)80097-1, 1968.
- 690 Bovensmann, H., Burrows, J. P., Buchwitz, M., Frerick, J., Noël, S., Rozanov, V. V., Chance, K. V., and Goede, A. P. H.: SCIAMACHY: Mission objectives and measurement modes, *J. Atmos. Sci.*, 56, 127–

- 150, doi:10.1175/1520-0469(1999)056<0127:SMOAMM>2.0.CO;2, 1999.
- Bramstedt, K., Noël, S., Bovensmann, H., Gottwald, M., and Burrows, J. P.: Precise pointing knowledge for SCIAMACHY solar occultation measurements, *Atmos. Meas. Tech.*, 5, 2867–2880, doi:10.5194/amt-5-2867-2012, 2012.
- 695 Brooke, J. S. A., Bernath, P. F., Western, C. M., Sneden, C., Afsar, M., Li, G., and Gordon, I. E.: Line strengths of rovibrational and rotational transitions in the $X^2\Pi$ ground state of OH, *J. Quant. Spectrosc. Ra.*, 168, 142-157, doi:10.1016/j.jqsrt.2015.07.021, 2016.
- Burkholder, J. B., Sander, S. P., Abbatt, J., Barker, J. R., Huie, R. E., Kolb, C. E., Kurylo, M. J., Orkin, 700 V. L., Wilmouth, D. M., and Wine, P. H.: Chemical Kinetics and Photochemical Data for Use in Atmospheric Studies, Evaluation No. 18, JPL Publication 15-10, Jet Propulsion Laboratory, Pasadena, <http://jpldataeval.jpl.nasa.gov>, 2015.
- Burt, K. D. and Sharma, R. D.: Near-resonant energy transfer from highly vibrationally excited OH to N_2 , *J. Chem. Phys.*, 128, 124311-1 to 124311-8, doi:10.1063/1.2884343, 2008.
- 705 Caridade, P. J. S. B., Sabin, J., Garridoz, J. D., and Varandas, A. J. C.: Dynamics of OH + O₂ vibrational relaxation processes, *Phys. Chem. Chem. Phys.*, 4, 4959-4969, doi:10.1039/b203101a, 2002.
- Caridade, P. J. S. B., Horta, J. Z. J., and Varandas, A. J. C.: Implications of the O + OH reaction in hydroxyl nightglow modeling, *Atmos. Chem. Phys.*, 13, 1-13, doi:10.5194/acp-13-1-2013, 2013.
- Chalamala, B. R. and Copeland, R. A.: Collision dynamics of OH($X^2\Pi$, $v=9$), *J. Chem. Phys.*, 99, 5807- 710 5811, doi:10.1063/1.465932, 1993.
- Charters, P. E., Macdonald, R. G., and Polanyi, J. C.: Formation of vibrationally excited OH by the reaction H+O₃, *Appl. Optics*, 10, 1747-1754, doi:10.1364/AO.10.001747, 1971.
- Dodd, J. A., Lipson, S. J., and Blumberg, W. A. M.: Formation and vibrational relaxation of OH($X^2\Pi_i, v$) by O₂ and CO₂, *J. Chem. Phys.*, 95, 5752-5762, doi:10.1063/1.461597, 1991.
- 715 Dyer, M. J., Knutsen, K., and Copeland, R. A.: Energy transfer in the ground state of OH: Measurements of OH($v=8,10,11$) removal, *J. Chem. Phys.*, 107, 7809-7815, doi:10.1063/1.475094, 1997.
- Finlayson-Pitts, B. J. and Kleindienst, T. E.: The reaction of hydrogen atoms with ozone as a source of vibrationally excited OH($X^2\Pi_i$) _{$v=9$} for kinetic studies, *J. Chem. Phys.*, 74, 5643-5658,

720 doi:10.1063/1.440928, 1981.

García-Comas, M., López-Puertas, M., Marshall, B. T., Wintersteiner, P. P., Funke, B., Bermejo-Pantaleón, D., Mertens, C. J., Remsberg, E. E., Gordley, L. L., Mlynczak, M. G., and Russell III, J. M.: Errors in Sounding of the Atmosphere using Broadband Emission Radiometry (SABER) kinetic temperature caused by non-local-thermodynamic-equilibrium model parameters, *J. Geophys. Res.*, 113, D24106, doi:10.1029/2008JD010105, 2008.

Good, R. E.: Determination of atomic oxygen density from rocket borne measurements of hydroxyl airglow, *Planet. Space Sci.*, 24, 389–395, doi:10.1016/0032-0633(76)90052-0, 1976.

Gottwald, M., Krieg, E., von Savigny, C., Noël, S., Bovensmann, H., and Bramstedt, K.: Determination of SCIAMACHY Line of Sight Misalignments, Proceedings of the Envisat Atmospheric Science Conference, ESA SP-636, Montreux, Switzerland, 23–27 April, 2007.

Grygalashvily, M., Sonnemann, G. R., Lübken, F.-J., Hartogh, P., and Berger, U.: Hydroxyl layer: Mean state and trends at midlatitudes, *J. Geophys. Res. Atmos.*, 119, 12,391–12,419, doi:10.1002/2014JD022094, 2014.

Howell, C. D., Michelangeli, D. V., Allen, M., Yung, Y. L., and Thomas, R. J.: SME observations of $O_2(^1\Delta_g)$ nightglow: An assessment of the chemical production mechanisms, *Planet. Space Sci.*, 38, 529–537, doi:10.1016/0032-0633(90)90145-G, 1990.

Kalogerakis, K. S., Smith, G. P., and Copeland, R. A.: Collisional removal of $OH(X^2\Pi, v = 9)$ by O , O_2 , O_3 , N_2 , and CO_2 , *J. Geophys. Res.*, 116, D20307, doi:10.1029/2011JD015734, 2011.

Kalogerakis, K. S., Matsiev, D., Sharma, R. D., and Wintersteiner, P. P.: Resolving the mesospheric nighttime 4.3 μm emission puzzle: Laboratory demonstration of new mechanism for $OH(v)$ relaxation, *Geophys. Res. Lett.*, 43, 8835–8843, doi:10.1002/2016GL069645, 2016.

Kaufmann, M., Lehmann, C., Hoffmann, L., Funke, B., Lopez-Puertas, M., von Savigny, C., and Riese, M.: Chemical heating rates derived from SCIAMACHY vibrationally excited OH limb emission spectra, *Adv. Space Res.*, 41, 1914–1920, doi:10.1016/j.asr.2007.07.045, 2008.

Klenerman, D. and Smith, I. W. M.: Infrared chemiluminescence studies using a SISAM spectrometer, *J. Chem. Soc., Farad. T. 2*, 83, 229–241, doi:10.1039/F29878300229, 1987.

Knutsen, K., Dyer, M. J., and Copeland, R. A.: Collisional removal of $OH(X^2\Pi, v=7)$ by O_2 , N_2 , CO_2 ,

- and N₂O, *J. Chem. Phys.*, 104, 5798-5802, doi:10.1063/1.471311, 1996.
- 750 Kulikov, M. Y., Belikov, M. V., Grygalashvyly, M., Sonnemann, G. R., Ermakova, T. S., Nechaev, A. A., and Feigin, A. M.: Nighttime ozone chemical equilibrium in the mesopause region. *J. Geophys. Res.-Atmos.*, 123, 1-15, doi:10.1002/2017JD026717, 2018.
- Lacoursiere, J., Dyer, M. J., and Copeland, R. A.: Temperature dependence of the collisional energy transfer of OH(v=10) between 220 and 310 K, *J. Chem. Phys.*, 118, 1661-1667, doi:10.1063/1.1530581, 2003.
- 755 López-Puertas, M., García-Comas, M., Funke, B., Picard, R. H., Winick, J. R., Wintersteiner, P. P., Mlynczak, M. G., Mertens, C. J., Russell III, J. M., and Gordley, L. L.: Evidence for an OH(v) excitation mechanism of CO₂ 4.3 μm nighttime emission from SABER/TIMED measurements, *J. Geophys. Res.*, 109, D09307, doi:10.1029/2003JD004383, 2004.
- 760 López-Puertas, M., García-Comas, M., Funke, B., Gardini, A., Stiller, G. P., von Clarmann, T., Glatthor, N., Laeng, A., Kaufmann, M., Sofieva, V. F., Froidevaux, L., Walker, K. A., and Shiotani, M.: MIPAS observations of ozone in the middle atmosphere, *Atmos. Meas. Tech.*, 11, 2187-2212, doi: 10.5194/amt-11-2187-2018, 2018.
- Meinel, A. B.: OH Emission Bands in the Spectrum of the Night Sky. II, *Astrophys. J.*, 112, 120-130, doi:10.1086/145321, 1950.
- 765 Mlynczak, M. G. and Solomon, S.: A detailed evaluation of the heating efficiency in the middle atmosphere, *J. Geophys. Res.*, 98, 10,517–10,541, doi:10.1029/93JD00315, 1993.
- Mlynczak, M. G., Martin-Torres, F. J., Johnson, D. G., Kratz, D. P., Traub, W. A., and Jucks, K.: Observations of the O(³P) fine structure line at 63 mm in the upper mesosphere and lower thermosphere, *J. Geophys. Res.*, 109, A12306, doi:10.1029/2004JA010595, 2004.
- 770 Mlynczak, M. G., Martin-Torres, F. J., Crowley, G., Kratz, D. P., Funke, B., Lu, G., Lopez-Puertas, M., Russell III, J. M., Kozyra, J., Mertens, C., Sharma, R., Gordley, L., Picard, R., Winick, J., and Paxton, L.: Energy transport in the thermosphere during the solar storms of April 2002, *J. Geophys. Res.*, 110, A12S25, doi:10.1029/2005JA011141, 2005.
- Mlynczak, M. G., Hunt, L. A., Mast, J. C., Marshall, B. T., Russell III, J. M., Smith, A. K., Siskind, D. E., Yee, J.-H., Mertens, C. J., Martin-Torres, F. J., Thompson, R. E., Drob, D. P., and Gordley L. L.:
- 775

- Atomic oxygen in the mesosphere and lower thermosphere derived from SABER: Algorithm theoretical basis and measurement uncertainty, *J. Geophys. Res.-Atmos.*, 118, 5724–5735, doi:10.1002/jgrd.50401, 2013a.
- 780 Mlynczak, M. G., Hunt, L. H., Mertens, C. J., Marshall, B. T., Russell III, J. M., López-Puertas, M., Smith, A. K., Siskind, D. E., Mast, J. C., Thompson, R. E., and Gordley, L. L.: Radiative constraints on the global annual mean atomic oxygen concentration in the mesopause region, *J. Geophys. Res.-Atmos.*, 118, 5796-5802, doi:10.1002/jgrd.50400, 2013b.
- 785 Mlynczak, M. G., Hunt, L. A., Marshall, B. T., Mertens, C. J., Marsh, D. R., Smith, A. K., Russell III, J. M., Siskind, D. E., and Gordley, L. L.: Atomic hydrogen in the mesopause region derived from SABER: Algorithm theoretical basis, measurement uncertainty, and results, *J. Geophys. Res.-Atmos.*, 119, 3516–3526, doi:10.1002/2013JD021263, 2014.
- Mlynczak, M. G., Hunt, L. A., Russell III, J. M., and Marshall, B. T.: Updated SABER Night Atomic Oxygen and Implications for SABER Ozone and Atomic Hydrogen, *Geophys. Res. Lett.*, 45, 5735-5741, doi:10.1029/2018GL077377, 2018.
- 790 Murtagh, D. P., Witt, G., Stegman, J., McDade, J. C., Llewellyn, E. J., Harris, F., and Greer, R. G. H.: An assessment of proposed $O(^1S)$ and $O_2(b^1\Sigma_g^+)$ nightglow excitation parameters, *Planet. Space Sci.*, 38, 43-53, doi:10.1016/0032-0633(90)90004-A, 1990.
- 795 Nelson Jr., D. D., Schiffman, A., Nesbitt, D. J., Orlando, J. J., and Burkholder, J. B.: H + O₃ Fourier-transform infrared emission and laser absorption studies of OH($X^2\Pi$) radical: An experimental dipole moment function and state-to-state Einstein A coefficients, *J. Chem. Phys.*, 93, 7003-7019, doi:10.1063/1.459476, 1990.
- Ohoyama, H., Kasai, T., Yoshimura, Y., Kimura, H., and Kuwata, K.: Initial distribution of vibration of the OH radicals produced in the $H+O_3 \rightarrow OH(X^2\Pi_{1/2,3/2})+O_2$ reaction: Chemiluminescence by a crossed beam technique, *Chem. Phys. Lett.*, 118, 263-266, doi:10.1016/0009-2614(85)85312-4, 1985.
- 800 Panka, P. A., Kutepov, A. A., Kalogerakis, K. S., Janches, D., Russell III, J. M., Rezac, L., Feofilov, A. G., Mlynczak, M. G., and Yigit, E.: Resolving the mesospheric nighttime 4.3 μm emission puzzle: comparing the CO₂(ν_3) and OH(ν) emission models, *Atmos. Chem. Phys.*, 17, 9751-9760, doi:10.5194/acp-17-9751-2017, 2017.

- Panka, P. A., Kutepov, A. A., Rezac, L., Kalogerakis, K. S., Feolilov, A. G., Marsh, D., Janches, D., and
805 Yigit, E.: Atomic Oxygen Retrieved From the SABER 2.0- and 1.6- μm Radiances Using New First-
Principles Nighttime OH(ν) Model, *Geophys. Res. Lett.*, 45, 5798-5803, doi:10.1029/2018GL077677,
2018.
- Rensberger, K. J., Jeffries, J. B., and Crosley, D. R.: Vibrational relaxation of OH($X^2\Pi_i, v=2$), *J. Chem.*
Phys., 90, 2174-2181, doi:10.1063/1.456671, 1989.
- 810 Rothman, L. S., Gordon, I. E., Barbe, A., Benner, D. C., Bernath, P. F., Birk, M., Boudon, V., Brown, L.
R., Campargue, A., Champion, J.-P., Chance, K., Coudert, L. H., Danaj, V., Devi, V. M., Fally, S., Flaud,
J.-M., Gamache, R. R., Goldman, A., Jacquemart, D., Kleiner, I., Lacome, N., Lafferty, W. J., Mandin,
J.-Y., Massie, S. T., Mikhailenko, S. N., Miller, C. E., Moazzen-Ahmadi, N., Naumenko, O. V., Nikitin,
A. V., Orphal, J., Perevalov, V. I., Perrin, A., Predoi-Cross, A., Rinsland, C. P., Rotger, M., Simeckova,
815 M., Smith, M. A. H., Sung, K., Tashkun, S. A., Tennyson, J., Toth, R. A., Vandaele, A. C., Vander
Auwera, J.: The HITRAN 2008 molecular spectroscopic database, *J. Quant. Spectrosc. Ra.*, 110, 533–
572, doi:10.1016/j.jqsrt.2009.02.013, 2009.
- Russell, J. P. and Lowe, R. P.: Atomic oxygen profiles (80–94 km) derived from Wind Imaging
Interferometer/Upper Atmospheric Research Satellite measurements of the hydroxyl airglow: 1.
820 Validation of technique, *J. Geophys. Res.*, 108, 4662, doi:10.1029/2003JD003454, 2003.
- Russell, J. P., Ward, W. E., Lowe, R. P., Roble, R. G., Shepherd, G. G., and Solheim, B.: Atomic oxygen
profiles (80 to 115 km) derived from Wind Imaging Interferometer/Upper Atmospheric Research
Satellite measurements of the hydroxyl and greenline airglow: Local time–latitude dependence, *J.*
Geophys. Res., 110, D15305, doi:10.1029/2004JD005570, 2005.
- 825 Russell III, J. M., Mlynczak, M. G., Gordley, L. L., Tansock, J., and Esplin, R.: An overview of the
SABER experiment and preliminary calibration results, in *Proceedings of the 44th Annual Meeting*,
Denver, Colorado, July 18–23, vol. 3756, pp. 277–288, SPIE, Bellingham, WA, 1999.
- Sander, R., Baumgaertner, A., Gromov, S., Harder, H., Jöckel, P., Kerkweg, A., Kubistin, D., Regelin,
E., Riede, H., Sandu, A., Taraborrelli, D., Tost, H., and Xie, Z.-Q.: The atmospheric chemistry model
830 CAABA/MECCA-3.0, *Geosci. Model Dev.*, 4, 373-380, doi:10.5194/gmd-4-373-2011, 2011.
- Shalashilin, D. V., Umanskii, S. Y., and Gershenson, Y. M.: Dynamics of vibrational energy exchange in

- collisions of OH and OD radicals with N₂. Application to the kinetics of OH-vibrational deactivation in the upper atmosphere, *Chem. Phys.*, 168, 315-325, doi:10.1016/0301-0104(92)87165-6, 1992.
- Shalashilin, D. V., Michtchenko, A. V., Umanskii, S. Y., and Gershenzon, Y. M.: Simulation of Effective
835 Vibrational-Translational Energy Exchange in Collisions of Vibrationally Excited OH with O₂ on the Model Potential Energy Surface. Can the Relaxation of OH(v) Be One-Quantum for Low and Multiquantum for High v?, *J. Phys., Chem.-US*, 99, 11627-11635, doi:10.1021/j100030a001, 1995.
- Sharma, R. D., Wintersteiner, P. P., and Kalogerakis, K. S.: A new mechanism for OH vibrational relaxation leading to enhanced CO₂ emissions in the nocturnal mesosphere, *Geophys. Res. Lett.*, 42,
840 4639-4647, doi:10.1002/2015GL063724, 2015.
- Sharp, W. E. and Kita, D.: In situ measurement of atomic hydrogen in the upper mesosphere, *J. Geophys. Res.-Atmos.*, 92, 4319-4324, doi:10.1029/JD092iD04p04319, 1987.
- Smith, A. K., Marsh, D. R., Mlynczak, M. G., and Mast, J. C.: Temporal variation of atomic oxygen in the upper mesosphere from SABER, *J. Geophys. Res.*, 115, D18309, doi:10.1029/2009JD013434, 2010.
- 845 Smith, A. K., Harvey, V. L., Mlynczak, M. G., Funke, B., García-Comas, M., Hervig, M., Kaufmann, M., Kyrölä, E., López-Puertas, M., McDade, I., Randall, C. E., Russell III, J. M., Sheese, P. E., Shiotani, M., Skinner, W. R., Suzuki, M., and Walker, K. A.: Satellite observations of ozone in the upper mesosphere, *J. Geophys. Res.-Atmos.*, 118, 5803-5821, doi:10.1002/jgrd.50445, 2013.
- Streit, G. E. and Johnston, H. S.: Reactions and quenching of vibrationally excited hydroxyl radicals, *J.*
850 *Chem. Phys.*, 64, 95-103, doi:10.1063/1.431917, 1976.
- Thiebaud, J. E., Copeland, R. A, and Kalogerakis, K. S.: Vibrational relaxation of OH(v=7) with O, O₂ and H, Abstract SA43A-1752, Fall Meeting, AGU, San Francisco, Calif, 2010.
- Turnbull, D. N. and Lowe, R. P.: New hydroxyl transition probabilities and their importance in airglow studies, *Planet. Space Sci.*, 37, 723-738, doi:10.1016/0032-0633(89)90042-1, 1989.
- 855 Van der Loo, M. P. J., and Groenenboom, G. C.: Theoretical transition probabilities for the OH Meinel system, *J. Chem. Phys.*, 126, 114413, doi: 10.1063/1.2646859, 2007.
- Varandas, A. J. C.: Reactive and non-reactive vibrational quenching in O + OH collisions, *Chem. Phys. Lett.*, 396, 182-190, doi:10.1016/j.cplett.2004.08.023, 2004.
- Von Savigny, C. and Lednyts'kyi, O.: On the relationship between atomic oxygen and vertical shifts

860 between OH Meinel bands originating from different vibrational levels, *Geophys. Res. Lett.*, 40, 5821-5825, doi:10.1002/2013GL058017, 2013.

Von Savigny, C., McDade, I. C., Eichmann, K.-U., and Burrows, J. P.: On the dependence of the OH* Meinel emission altitude on vibrational level: SCIAMACHY observations and model results, *Atmos. Chem. Phys.*, 12, 8813–8828, doi:10.5194/acp-12-8813-2012, 2012.

865 Xu, J., Gao, H., Smith, A. K., and Zhu, Y.: Using TIMED/SABER nightglow observations to investigate hydroxyl emission mechanisms in the mesopause region, *J. Geophys. Res.*, 117, D02301, doi:10.1029/2011JD016342, 2012.

Zhu, Y., and Kaufmann, M.: Atomic oxygen abundance retrieved from SCIAMACHY hydroxyl nightglow measurements, *Geophys. Res. Lett.*, 45, 1-9, doi:10.1029/2018GL079259, 2018.

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Table 1. Physical processes and chemical reactions included in the Base model

	Process	Rate or scheme	Reference
R1	$\text{H} + \text{O}_3 \rightarrow \text{OH}(\nu) + \text{O}_2$	$k_1 = 1.4 \times 10^{-10} e^{-(470/T)}$ $k_1(\nu) = k_1 f_1(\nu)^a$	Burkholder et al. (2015), Alder-Golden (1997, Table 1)
R2	$\text{O}({}^3\text{P}) + \text{O}_3 \rightarrow \text{O}_2 + \text{O}_2$	$k_2 = 8 \times 10^{-12} e^{-(2060/T)}$	Burkholder et al. (2015)
R3	$\text{O}({}^3\text{P}) + \text{O}_2 + \text{M} \rightarrow \text{O}_3 + \text{M}$	$k_3 = 6 \times 10^{-34} (300/T)^{2.4}$	Burkholder et al. (2015)
R4	$\text{OH}(\nu) \rightarrow \text{OH}(\nu') + h\nu$	variable rates	Xu et al. (2012, Table A1)
R5	$\text{OH}(\nu) + \text{N}_2 \rightarrow \text{OH}(\nu') + \text{N}_2$	$\nu' = \nu - 1$	Adler-Golden (1997, Table 1), Kalogerakis et al. (2011)
R6	$\text{OH}(\nu) + \text{O}_2 \rightarrow \text{OH}(\nu') + \text{O}_2$	$\nu' < \nu$	Adler-Golden (1997, Table 3), see text for more information
R7a	$\text{OH}(\nu) + \text{O}({}^3\text{P}) \rightarrow \text{H} + \text{O}_2$	variable rates	Varandas (2004, Table 3, M I)
R7b	$\text{OH}(\nu) + \text{O}({}^3\text{P}) \rightarrow \text{OH}(\nu') + \text{O}({}^3\text{P})$	$\nu' < \nu$	Caridade et al. (2013, Table 1)

Rate constants are given in the $\text{cm}^3\text{-s}^{-1}$ system.

890 ^a $f_1(5, 6, 7, 8, 9) = 0.01, 0.03, 0.15, 0.34, 0.47$

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Table 2. Empirically determined branching ratios of $\text{OH}(v)+\text{O}_2\rightarrow\text{OH}(v')+\text{O}_2$ of the O_2 best fit model based on $\text{OH}(6-2)$ VER, $\text{OH}(5-3)+\text{OH}(4-2)$ VER, and $\text{OH}(3-1)$ VER observations.

v/v'	8	7	6	5	4	3	≤ 2
9	0	0	0	0	1	0	0
8		0	0	0	0.3	0.7	0
7			0	0	0	0.1	0.9
6				0	0	0	1
5					0	0	1
4						0	1
3							1

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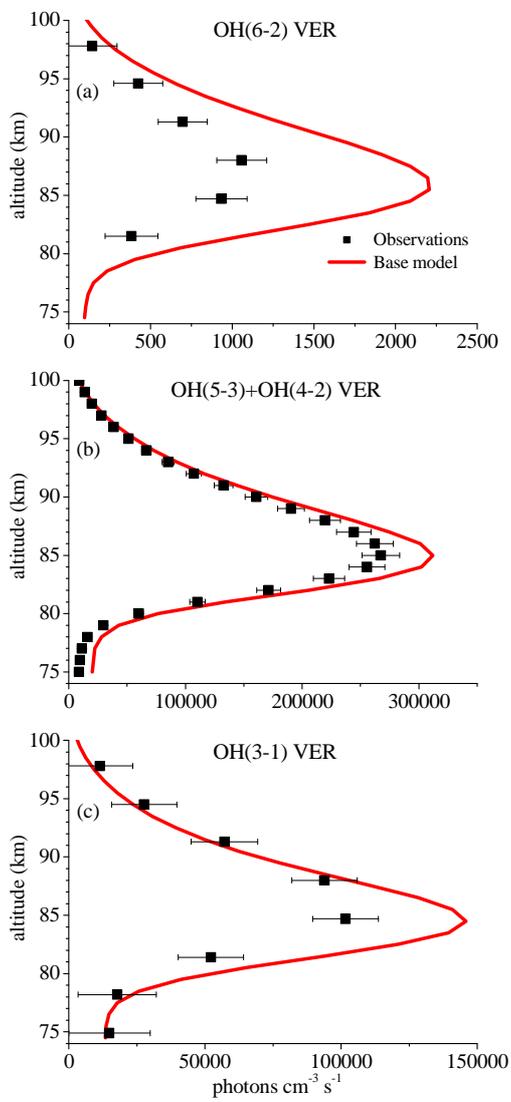
Table 3. Empirically determined branching ratios of $\text{OH}(v)+\text{O}(^3\text{P})\rightarrow\text{OH}(v')+\text{O}(^1\text{D})$ of the Best fit model based on OH(6-2) VER, OH(5-3)+OH(4-2) VER, and OH(3-1) VER observations.

	Process	Recommendation	Best fit rate ($\text{cm}^3 \text{s}^{-1}$)
R11a	$\text{OH}(9) + \text{O}(^3\text{P}) \rightarrow \text{OH}(4) + \text{O}(^1\text{D})$	$k_{11}(9-4) > 0.6 \times k_{11}(v=9)$	$0.8 \times 2.30 \times 10^{-10}$
R11b	$\text{OH}(9) + \text{O}(^3\text{P}) \rightarrow \text{OH}(3) + \text{O}(^1\text{D})$	not negligible	$0.2 \times 2.30 \times 10^{-10}$
R11c	$\text{OH}(8) + \text{O}(^3\text{P}) \rightarrow \text{OH}(3) + \text{O}(^1\text{D})$	$k_{11}(v=8) < k_{11}(v=9)$	$1.0 \times 1.80 \times 10^{-10}$
R11d	$\text{OH}(7) + \text{O}(^3\text{P}) \rightarrow \text{OH}(\leq 2) + \text{O}(^1\text{D})$	$k_{11}(7-\leq 2) < k_{11}(v=8)$	1.25×10^{-10}
R11fe	$\text{OH}(6) + \text{O}(^3\text{P}) \rightarrow \text{OH}(\leq 1) + \text{O}(^1\text{D})$	$k_{11}(6-\leq 1) < k_{11}(v=7)$	0.80×10^{-10}
R11gf	$\text{OH}(5) + \text{O}(^3\text{P}) \rightarrow \text{OH} + \text{O}(^1\text{D})$	$k_{11}(v=5) < k_{11}(v=6)$	0.40×10^{-10}

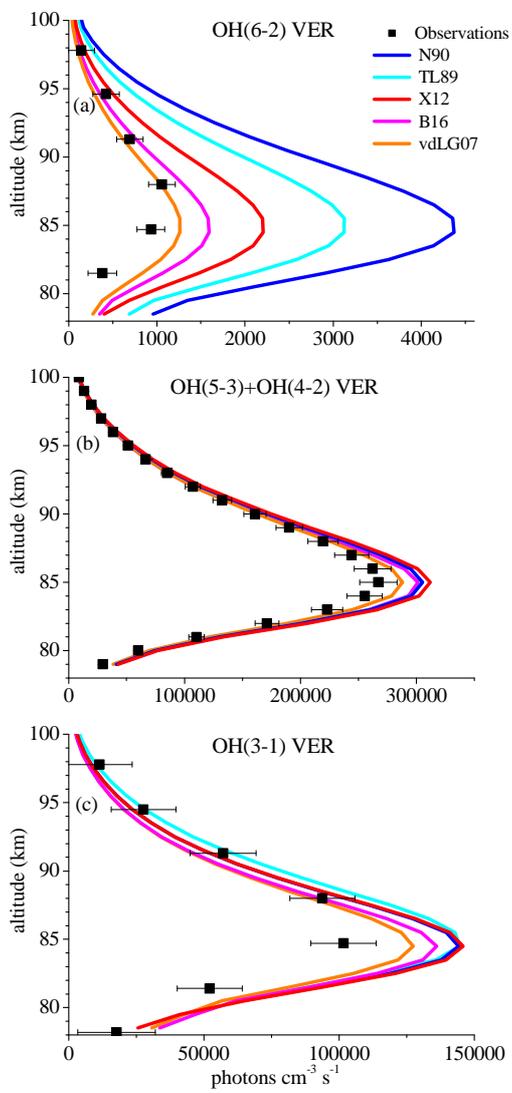
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950 **Figure 1 :** Comparison of vertical profiles of the volume emission rate (VER) of a) OH(6-2), b) OH(5-3)+OH(4-2), and c) OH(3-1) at 0°-10° N between satellite observations and the Base model output. The observations are climatology of night-time mean zonal means from 2003 to 2011, based on co-location measurements of TIMED/SABER and ENVISAT/SCIAMACHY. Note the different scaling of the x-axis.



955 **Figure 2** : Same as Figure 1 but for different sets of Einstein coefficients from literature, namely N90 (Nelson et al., 1990), TL89 (Turnbull and Lowe, 1989), X12 (=Base model; Xu et al., 2012), B16 (Brooke et al., 2016), and vdLG07 (van der Loo and Groenenboom, 2007).

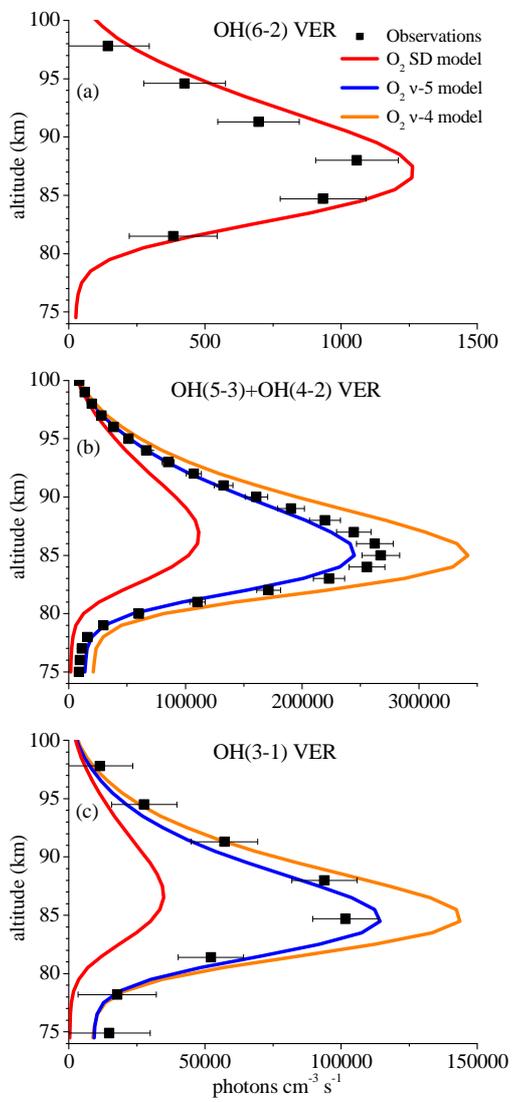


Figure 3 : Same as Fig. 1 but for the O₂ SD model, the O₂ v-5 model, and the O₂ v-4 model. Note that the results of these three models are identical in case of OH(6-2) VER.

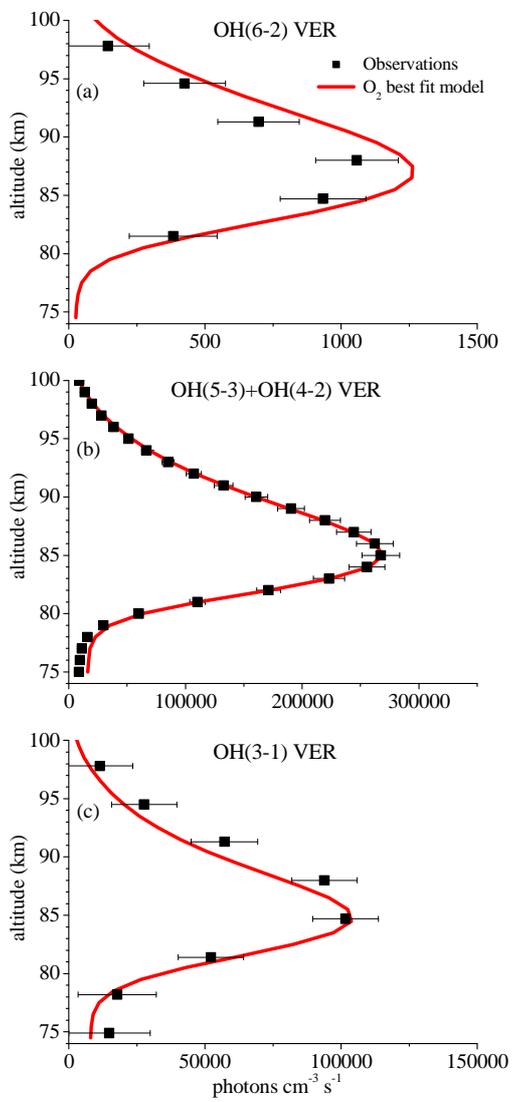


Figure 4 : Same as Fig. 1 but for the O_2 best fit model. Note that Fig. 4a is identical to Fig. 3a but was plotted again for convenience.

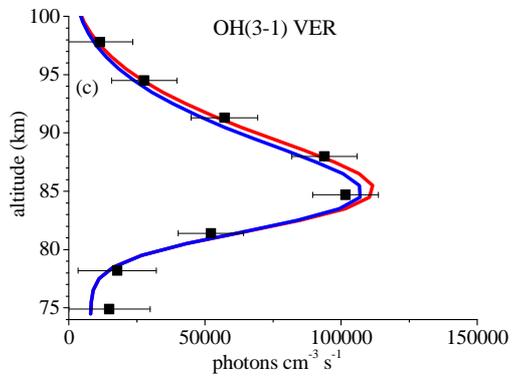
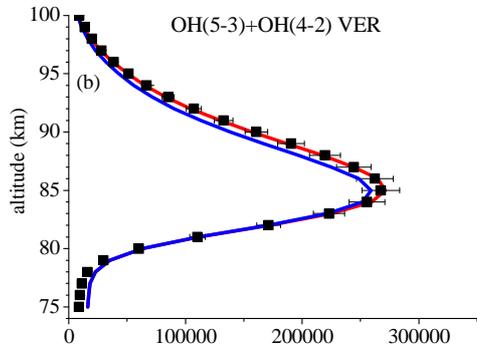
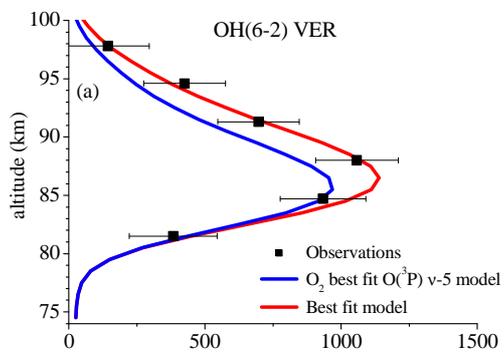


Figure 5 : Same as Fig. 1 but for the O_2 best fit $O(^3P)$ v-5 model and the Best fit model.

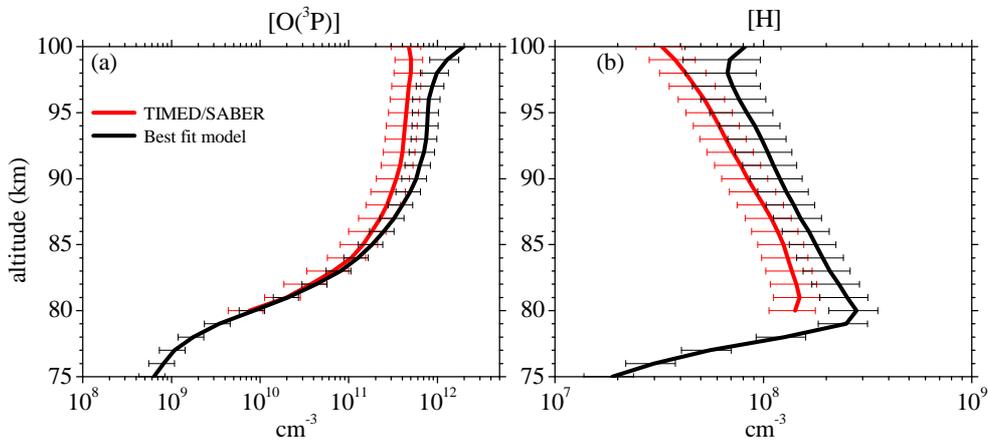


Figure 6 : Vertical profiles of a) $[O(^3P)]$ and b) $[H]$ derived from SABER OH(9-7)+OH(8-6) VER observations (Mlynczak et al., 2018) and our Best fit model by fitting SABER OH(9-7)+OH(8-6) VER and OH(5-3)+OH(4-2) VER as well as SCIAMACHY OH(6-2) VER and OH(3-1) VER. Shown are averages of night-time mean zonal means of co-location measurements (see Sect. 2.2)

975 from 2003 to 2011 between 0° and 10° N. Error bars show the 1σ uncertainty due to chemical and physical processes.

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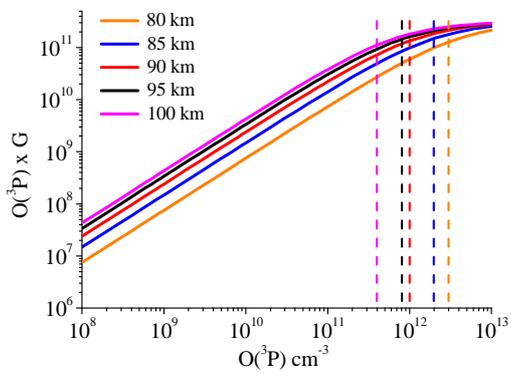


Figure 7 : $O(^3P) \times G$ as a function of $O(^3P)$ at different altitudes. The visually determined upper limits of $O(^3P)$ before non-linearity becomes too pronounced are represented by the dashed lines.