Atmospheric Band Fitting Coefficients Derived from Self-Consistent Rocket-Borne Experiment

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Abstract

Based on self-consistent rocket-borne measurements of temperature, densities of atomic oxygen and neutral air, and volume emission of the Atmospheric Band (762 nm) we examined the one-step and two-step excitation mechanism of $O_2\left(b^1\Sigma_g^+\right)$ for night-time conditions. Following McDade et al. (1986), we derived the empirical fitting coefficients, which parameterize the Atmospheric Band emission $O_2\left(b^1\Sigma_g^+ - X^3\Sigma_g^-\right)(0,0)$. This allows to derive atomic oxygen concentration from night-time observations of Atmospheric Band emission $O_2\left(b^1\Sigma_g^+ - X^3\Sigma_g^-\right)(0,0)$. The derived empirical parameters can also be utilised for Atmospheric Band modelling. Additionally, we derived fit function and corresponding coefficients for combined (one- and two-step) mechanism. Simultaneous common volume

measurements of all the parameters involved in the theoretical calculation of the observed $O_2(b^1\Sigma_g^+ - X^3\Sigma_g^-)(0,0)$ emission, i.e. temperature and density of the background air, atomic oxygen density, and volume emission rate, is the novelty and the advantage of this work.

1. Introduction

The mesopause region is essential to understand the chemical and physical processes in the upper atmosphere because this is region of coldest temperature (during summer at high latitudes) and highest turbulence in the atmosphere (e.g. Lübken, 1997), the region of formation of such phenomena as NLC and PMSE (e.g. Rapp and Lübken, 2004), the region of gravity waves (GWs) breaking and formation of secondary GWs (Becker and Vadas, 2018), as well as the region of coupling between mesosphere and thermosphere. This region is characterised by different airglow emissions and, particularly, by the emissions of the Atmospheric Band which is produced by the excited state of molecular oxygen $O_2(b^1\Sigma_g^+)$. Airglow observation in the Atmospheric Band is a useful method to study dynamical processes in the mesopause region. There have been a number of reports of gravity waves (GWs) detection in this band (Noxon, 1978; Viereck and Deehr, 1989; Zhang et al., 1993). Planetary wave climatology has been investigated by the Spectral Airglow Temperature Imager (SATI) instrument (Lopez-Gonzalez et al., 2009). In addition, the parameters of tides have been reported from SATI (Lopez-Gonzalez et al., 2005) and High Resolution Doppler Imager (HRDI) observations (Marsh et al., 1999). In number of works Sheese et al. (2010, 2011) inferred the temperature by Atmospheric Band observation. Furthermore, the response of mesopause temperature and atomic oxygen during major sudden stratospheric warming was studied utilising Atmospheric Band emission by Shepherd et al. (2010). Various works have focused on Atmospheric Band emission modelling with respect to gravity waves and tides (e.g. Hickey et al., 1993; Leko et al., 2002; Liu and Swenson, 2003). The specific theory

of the gravity wave effects on $O_2\left(b^1\Sigma_g^+\right)$ emission was derived in Tarasick and Shepherd (1992). Moreover, Atmospheric Band observations have been widely utilised to infer atomic oxygen, which is an essential chemical constituent for energetic balance in the extended mesopause region (e.g. Hedin et al., 2009, and references there in), and ozone concentration (Mlynczak et al., 2001). Although there is a large field of application of Atmospheric Band emissions, there is a lack of knowledge on processes of the $O_2\left(b^1\Sigma_g^+\right)$ population. Two main mechanisms of night-time population (note, the day-time mechanisms are quite different, see e.g. Zarboo et al., (2018)) were proposed: the first is the direct population from three-body recombination of atomic oxygen (e. g. Deans et al., 1976); the second is the so-called two-step mechanism, which assumes an intermediate excited precursor O_2^* (e. g. Witt et al., 1984; Greer et al., 1981). It has been shown by laboratory experiments that the first mechanism alone has not explained observed emissions (Young and Sharpless, 1963; Clyne at al., 1965; Young and Black, 1966; Bates, 1988). The second mechanism entails a discussion about the precursor excited state and additional ambiguities in their parameters (e.g. Greer et al., 1981; Ogryzlo et al., 1984). Thus, Witt et al. (1984) proposed the hypothesis that the $O_2(c^1\Sigma_u^-)$ state is, possibly, the precursor; López-González et al. (1992a) suppose that the precursor could be $O_2(^5\Pi_g)$; Wildt et al. (1991) found by laboratory measurements that it could be $O_2(A^3\Sigma_u^+)$. Hence, the problem of identification is still not solved. The essential step in this direction has been done after the ETON 2 (Energy Transfer in the Oxygen Nightglow) rocket experiment. ETON 2 mission yielded empirical fitting parameters that allow either to quantify the $O_2\left(b^1\Sigma_g^+\right)$ (and, consequently, volume emission) by known O, or atomic oxygen by known volume emission values (McDade et al., 1986). Despite the significance of this work, the temperature and density of air (necessary for derivation) were taken from CIRA-72 and MSIS-83 (Hedin, 1983) models. This leads to some degree of uncertainty (e.g. Murtagh et al., 1990). Thus, more solid knowledge on these fitting coefficients based on consistent

measurements of atomic oxygen, volume emission of Atmospheric Band, and temperature and density of background atmosphere is desirable. In this paper we present common volume measurements of these parameters performed in the course of WADIS-2 sounding rocket mission. In the next section, we describe the rocket experiment and obtained data relevant for our study. In section 3, to make the paper easier to understand, we repeat some theoretical approximations from McDade et al. (1986). The obtained results of our calculations are discussed in section 4. Concluding remarks and summary are given in the last section.

2. Rocket experiment Description

The WADIS (Wave propagation and dissipation in the middle atmosphere: Energy budget and distribution of trace constituents) sounding rocket mission aimed to simultaneously study the propagation and dissipation of GWs and measure the concentration of atomic oxygen. It comprised two field campaigns conducted at the Andøya Space Center (ASC) in northern Norway (69°N, 16°E). The WADIS-2 sounding rocket was launched during the second campaign on 5 March 2015 at 01:44:00 UTC, that is, night-time conditions. For a more detailed mission description, the reader is referred to Strelnikov et al. (2017) and the accompanying paper by Strelnikov et al. (2018).

The WADIS-2 sounding rocket was equipped with the CONE instrument to measure absolute neutral air density and temperature with high spatial resolution, instrument for atomic oxygen density measurements FIPEX (Flux Probe Experiment) and the Airglow Photometer for atmospheric band (762 nm) volume emission observation.

CONE (COmbined measurement of Neutrals and Electrons), operated by IAP (Leibniz Institute of Atmospheric Physics at the Rostock University), is a classical triode type ionisation gauge optimised for a pressure range between 10^{-5} to 1 mbar. The triode system is surrounded by two electrodes: Whilst the outermost grid is biased to +3 to +6 V to measure

electron densities at a high spatial resolution, the next inner grid (-15 V) is meant to shield the ionisation gauge from ionospheric plasma. CONE is suitable for measureing absolute neutral air number densities at altitude range between 70 and 120 km. To obtain absolute densities, the gauges are calibrated in the laboratory using a high-quality pressure sensor, like a Baratron. The measured density profile can be further converted to a temperature profile assuming hydrostatic equilibrium. For a detailed description of the CONE instrument, see Giebeler et al. (1993) and Strelnikov et al. (2013). Molecular oxygen and molecular nitrogen are derived from CONE atmospheric number density measurements and partitioning according to NRLMSISE-00 reference atmosphere (Picone et al., 2002).

The Airglow Photometer operated by MISU (Stockholm University, Department of Meteorology) measures the emission of the molecular oxygen Atmospheric Band around 762 nm from the overhead column, from which volume emission rate is inferred by differentiation. For airglow measurements in general, a filter photometer is positioned under the nose cone viewing along the rocket axis with a defined field-of-view (FOV). For WADIS-2 however, the FOV of ±3° was tilted from the rocket axis by 3° to avoid having other parts of the payload within the FOV and at the same time roughly view the same volume as the other instruments. The optical design is a standard radiometer-type system with an objective lens, a field lens, aperture and stops which provide an even illumination over a large portion of the detector surface (photomultiplier tube) and a defined FOV. At the entrance of the photometer there is an interference filter with a passband of 6 nm centred at 762 nm. During ascent, after the nosecone ejection, the photometer then counts the incoming photons from the overhead column (or actually the overhead cone). When the rocket passes through the layer the measured photon flux drops and above the emission layer only weak background emissions are present (e.g. the zodiacal and galactic light). After the profile has been corrected for background emissions and attitude (van Rhijn effect) it is converted from counts to radiance using pre-flight laboratory calibrations. The calibration considers the spectral shape of the 0-0

band of the $O_2(b^1\Sigma_g^+ - X^3\Sigma_g^-)(0,0)$ Atmospheric Band system and the overlap of the interference filter passband. The profile is then smoothed and numerically differentiated with respect to altitude to yield the volume emission rate of the emitting layer. The data were sampled with 1085 Hz which results in an altitude resolution of about 0.75 m during the passage of the airglow layer (the velocity was ~800m/s at 95 km). However, because of the high noise level, the profile needed to be averaged to a vertical resolution of at least 3 km in order to get satisfactory results after the differentiation. A more detailed description and review of this measurement technique is given by Hedin et al. (2009).

The aim of the FIPEX developed by the IRS (Institute of Space Systems, University of Stuttgart) is to measure the atomic oxygen density along the rocket trajectory with high spatial resolution. It employs solid electrolyte sensor which has a selective sensitivity to atomic oxygen. A low voltage is applied between anode and cathode pumping oxygen ions through the electrolyte ceramic (yttria stabilised zirconia). The current measured is proportional to the oxygen density. Sampling is realised with a frequency of 100 Hz and enables a spatial resolution of ~10 m. Laboratory calibrations were done for molecular and atomic oxygen. For a detailed description of the FIPEX instruments and their calibration techniques see Eberhart et al. (2015, 2018).

3. Theory

Here, we are repeating the theory of $O_2(b^1\Sigma_g^+ - X^3\Sigma_g^-)(0,0)$ night-time emissions following McDade et al. (1986) to make our paper more readable, using all nomenclature as in the original paper. All utilised reactions are listed in Table 1, together with corresponding reaction rates, branching ratios, quenching rates and spontaneous emission coefficients. Some components have been updated according to modern knowledge, thus, deviating from the work of McDade et al. (1986).

Assuming direct one-step mechanism as a main one for population and that the $O_2\left(b^1\Sigma_g^+\right)$ in photochemical equilibrium, we can write its concentration as a ratio of production to the loss term:

$$\left[O_2\left(b^1\Sigma_g^+\right)\right] = \frac{\varepsilon k_1 [O]^2 M}{A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]}, \quad (1)$$

where k_1 – reaction rate for three-body recombination of atomic oxygen, ε is the corresponding quantum yield of $O_2\left(b^1\Sigma_g^+\right)$ formation, A_2 represents the spontaneous emission coefficient, and $k_2^{O_2}, k_2^{N_2}, k_2^{O}$ are the quenching coefficients for reactions with O_2, N_2 and O, respectively.

In case of known temperature, volume emission and concentrations of O, O_2, N_2, and M, the quantum yield of $O_2\left(b^1\Sigma_g^+\right)$ formation can be calculated as follows:

$$\varepsilon = V_{at} \frac{A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]}{A_1 k_1 [O]^2 M}, \quad (2)$$

where A_1 is spontaneous emission for reaction R5 (hereafter, nomenclature RX means the reaction X from Table 1).

In the case of the two-step mechanism, the unknown excited state O_2^* is populated at the first step from the reaction R7. Then, it can be deactivated by quenching (R9), spontaneous emission (R10) or producing $O_2\left(b^1\Sigma_g^+\right)$ by the reaction R8. Note, R8 is one pathway of the overall quenching reaction R9.

In the second step, O_2^* is transformed into $O_2\left(b^1\Sigma_g^+\right)$, which, can be deactivated by quenching (R2-R4) and by spontaneous emission (R6). Assuming photochemical equilibrium for O_2^* and, as before, for $O_2\left(b^1\Sigma_g^+\right)$ the volume emission in the case of $O_2\left(b^1\Sigma_g^+ - X^3\Sigma_g^-\right)(0,0)$ is:

$$V_{at} = \frac{A_1 \alpha k_1 [O]^2 M \gamma k_3^{O_2}[O_2]}{\left(A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]\right)\left(A_3 + k_3^{O_2}[O_2] + k_3^{N_2}[N_2] + k_3^{O}[O]\right)}, \quad (3)$$

where quantum yield of O_2^* formation α, quantum yield of $O_2\left(b^1\Sigma_g^+\right)$ formation γ, spontaneous emission coefficient A_3, and $k_3^{O_2}, k_3^{N_2}, k_3^O$ unknown quenching rates of O_2^*. Note, assumption about photochemical equilibrium for O_2^* and $O_2\left(b^1\Sigma_g^+\right)$ is valid, because $O_2\left(b^1\Sigma_g^+\right)$ radiative lifetime is less than 12 s and all potential candidates for of O_2^* have lifetime less than several seconds (e.g. López-González et al., 1992a, 1992b, 1992c; Yankovsky et al., 2016, and references therein).

Collecting all known values on the right-hand side (RHS) and all unknown summands on the left-hand side (LHS), omitting emissive summand A_3 as non-effective loss (McDade et al., 1986) equation (3) can be rearranged as follows:

$$C^{O_2}[O_2] + C^O[O] = \frac{A_1 k_1 [O]^2 M[O_2]}{V_{at}\left(A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^O[O]\right)}, \quad (4)$$

where $C^{O_2} = \left(1 + k_3^{N_2}[N_2]/k_3^{O_2}[O_2]\right)/\alpha\gamma$ and $C^O = k_3^O/\alpha\gamma k_3^{O_2}$ are the fitting coefficients that can be calculated by the least square fit (LSF) procedure. Such derivation assumes that the coefficients are temperature independent (or temperature dependence is weak). This means that the reaction rates k_3 are assumed to be temperature independent (dependence is weak), or have the same temperature dependency for all quenching partners (N_2, O_2, O). Currently, this statement on the basis of available information about potential precursors is submitted true, but for which the solid evidence is absent. We calculated them based on our measurements and will discuss the results in the following section.

In more general case population of $O_2\left(b^1\Sigma_g^+\right)$ occurs via both channels: one-step and two-step. We discuss such ability in section 4.3 and derive an expression for corresponding fit-function in Appendix.

4. Results and Discussion

Figure 1 shows input data for our calculations: temperature from CONE instrument (Fig. 1a), number density of air (Fig. 1b), atomic oxygen concentration measured by FIPEX (Fig. 1c) and volume emission at 762 nm from photometric instrument (Fig. 1d). A temperature minimum of ~158 K was observed at 104.2 km. A local temperature peak was measured at 98.9 km with values of 204.5 K. The secondary temperature minimum was visible at 95.4 km and amounted to ~173 K. Atomic oxygen concentration (Fig. 1c) had a peak of ~$4.7 \cdot 10^{11}$ [cm^{-3}] at 97.2 km and approximately coincided with the secondary temperature peak. The peak of volume emission was detected between 95 and 97 km with values of more than 1700 [$phot. \cdot cm^{-3} \cdot s^{-1}$]; this is slightly beneath the atomic oxygen corresponding maximum and slightly above the secondary temperature minimum. Note, this point to the competition of the temperature and the atomic oxygen concentration in the processes of atomic oxygen excited state $O_2\left(b^1\Sigma_g^+\right)$ formation. Independently of the mechanism of atmospheric band emission (Eq. 1 or Eq. 3), the numerator is directly proportional to the square of atomic oxygen concentration and inversely proportional to the third power of the temperature (via reaction rate k_1 and M, considering the ideal gas low). Our rocket experiment shows an essential difference of emissions between ascending and descending flights (see Strelnikov et al., 2018). It also demonstrates a significant variability in other measured parameters, including neutral temperature and density as well as atomic oxygen density (Strelnikov et al., 2017, 2018). This suggests that, in the case of the ETON 2 experiments, the temporal extrapolation of atomic oxygen for the time of the emission measurement flight (which was approximately 20 min earlier) may lead to serious biases in estimations because, as one can see from Eq. 1 and Eq. 3, volume emission depends on the atomic oxygen concentration quadratically. Since the best quality data were obtained during the descent of the WADIS-2 rocket flight, we chose this data set for our analysis (Strelnikov et al., 2018). The region above 104 km is subject to auroral contamination. In the region below 92 km, negative values may occur in the volume emission profile as the result of self-absorption in the denser atmosphere below the emission

layer. Hence, we considered the region near the peak of emission between 92 km and 104 km as most appropriate for our study. The comparisons of our measurements with other observations, as well as with the results of modelling are presented in several papers (e.g. Eberhart et al., 2018; Strelnikov et al., 2018).

4.1 One-step mechanism

Figure 2 shows the quantum yield of $O_2\left(b^1\Sigma_g^+\right)$ formation ε calculated according to Eq. (2), which is necessary to form $O_2\left(b^1\Sigma_g^+\right)$ under the assumption that the direct three-body recombination of atomic oxygen is the main mechanism. The uncertainties for this figure (as well as for other figures) were calculated according with sensitivity analysis (von Clarmann, 2014; Yankovsky and Manuilova, 2018, (Appendix 1); Vorobeva et al., 2018). Calculated values of ε are placed in the range [0.07; 0.13], which is in good agreement with the values derived by McDade et al. (1986). The averaged value amounts to 0.11±0.018. The range of values taking into account both the variance and the error range amounts to [0.02; 0.22]. By the physical nature of this value, the quantum yield of $O_2\left(b^1\Sigma_g^+\right)$ formation should not depend on altitude. Fig. 2 shows some altitude dependence of central values of ε, but considering the large error range, there is no clear altitude dependence. The variability of the data points is much smaller than the errors of the individual points. Hence, in light of the analysis of our rocket experiment, we may not decline direct excitation mechanism.

Although the population via one-step mechanism alone is, generally speaking, possible, it is improbable because laboratory experiments show that the direct excitation alone may not explain observed emissions (Young and Sharpless, 1963; Clyne at al., 1965; Young and Black, 1966; Bates, 1988). This conclusion is in agreement with the conclusion from McDade et al. (1986), which stated that the one-step excitation mechanism is not sufficient to explain

247 the $O_2(b^1\Sigma_g^+)$ population. Hence, in the following, we check the second energy transfer
248 mechanism.
249
250 **4.2 Two-step mechanism**
251
252 Figure 3 depicts the altitude profile of the right hand side (RHS) of equation (4) and profile
253 calculated by the least-square fit (LSF). The fitting coefficients, C^{O2} and C^{O}, resulting from
254 this fit, amount to $9.8^{-5.3}_{+6.5}$ and $2.1^{+0.3}_{-0.6}$, respectively. The uncertainties were calculated, as
255 commonly for LSF (Bevington and Robinson, 2003), based on error propagation from the
256 error and variance of RHS as provided in Figure 3. Our C^{O2} coefficient is partially, within the
257 error range, in agreement with C^{O2} coefficients given in McDade et al. (1986), which amount
258 to 4.8±0.3 and 6.6±0.4 for temperature from CIRA-72 and MSIS-83, respectively. The C^{O}
259 coefficient is approximately one order lower. There are several possible reasons for the large
260 discrepancy in C^{O}, for example the temperature dependence of the O-quenching or that, in the
261 case of ETON 2 experiments mean temperature profiles from the models CIRA-72 and MSIS-
262 83 were utilized, which does not reproduce any short-time dynamical fluctuations, solar cycle
263 conditions, etc. In frame of our analysis, we may not identify the reason for the large
264 discrepancy in C^{O} more precisely. Fitting coefficients defined in such a way (Eq. 4) do not
265 have a direct physical meaning. However, they have a physical meaning in several limit cases.
266 If the quenching coefficients of a precursor with molecular nitrogen are much smaller than
267 those with molecular oxygen $(k_3^{N_2} \ll k_3^{O_2})$, then $\alpha\gamma = 1/C^{O2}$. The assumption, that the
268 quenching of the precursor with N_2 is much slower than quenching with O_2, is just working
269 hypothesis, which is commonly used for analysis of possible precursor (e.g. McDade et al.,
270 1986; López-González et al, 1992a, 1992b; and references therein). It is true for such potential
271 precursor as $O_2(A^3\Sigma_u^+)$ (Kenner and Ogryzlo, 1983b), but generally, there is no evidence,
272 neither for nor against that. If it is not true, any definite conclusion on precursor by known

C^{O2} is not possible. In our case $\alpha\gamma = 0.102^{+0.120}_{-0.041}$. In other words, in the case of two-step formation of $O_2\left(b^1\Sigma_g^+\right)$ with energy transfer agent O_2, the total efficiency $\eta = \alpha\gamma$ amounts to 10.2%, which is the lowest amongst known values. Based on rocket experiment data analysis (ETON), Witt et al. (1984) obtained $\alpha\gamma = 0.12 - 0.2$. According to McDade et al. (1986), for the case with $k_2^O = 8 \cdot 10^{-14}$, the total efficiencies are 0.15 and 0.21 for temperature profiles adopted from MSIS-83 and CIRA-72, respectively. The analyses of López-González et al. (1992a, c), adopted O_2, N_2, and temperature profiles from the model (Rodrigo et al., 1991), showed a total efficiency of 0.16. In contrast to our work, all investigations mentioned above utilised the temperature and atmospheric density from models which describe a mean state of the atmosphere. This is a possible reason for discrepancy in the results. Total efficiency η may serve as an auxiliary quantity to identify the precursor. According to the physical meaning of efficiency, it may not be larger than 1. Hence, α, γ, as well as the total efficiency are smaller than 1. Consequently, $\gamma = \eta/\alpha < 1$, and we can examine potential candidates for O_2^* with this criterion. From an energetic point of view, only four bound states of molecular oxygen can be considered as an intermediate state for the $O_2\left(b^1\Sigma_g^+\right)$ population: $O_2(A^3\Sigma_u^+), O_2(A'^3\Delta_u), O_2(c^1\Sigma_u^-)$, and $O_2(^5\Pi_g)$ (Greer et al., 1981; Wraight, 1982; Witt et al., 1984; McDade et al., 1986; López-González et al., 1992c). For better readability, we will partially repeat a table from López-González et al. (1992b, c) with known α in our work (Table 2). From Table 2, it can be seen that only $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$ fit to the criterion of $\gamma = 0.102/\alpha < 1$. At lower limit of uncertainty ($\gamma = 0.061/\alpha < 1$) $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$ satisfy to the criterion, and considering upper limit ($\gamma = 0.222/\alpha < 1$), only $O_2(^5\Pi_g)$ may serve as precursor.

The second expression that helps to clarify the choice of the precursor is the ratio of quenching rates. In the limit of low quenching with molecular nitrogen $\left(k_3^{N_2} \ll k_3^{O_2}\right)$, the ratio of fitting coefficients equals the ratio of the quenching rates of atomic and molecular oxygens

$\left(C^O/C^{O2} = k_3^O/k_3^{O_2}\right)$. An analysis from the ETON 2 rocket experiment yields values of quenching coefficients ratios of potential precursor of 3.1 and 2.9 for temperatures from CIRA-72 and MSIS-83, respectively. This is close to the value from Ogryzlo et al. (1984), who found $k_3^O/k_3^{O_2} = 2.6$ by laboratory measurements; however, as was noted in their work, substitution of these values into the equation for emission yields 16 % of the observed emission (Ogryzlo et al., 1984). These findings point to the possibility of a too high measured ratio $k_3^O/k_3^{O_2}$ as the result of too strong quenching of precursor by atomic oxygen. Our value of quenching ratios $k_3^O/k_3^{O_2}$ amounts to $0.21^{+0.32}_{-0.12}$. There is not enough information on measured values for bound states of molecular oxygen. Laboratory measurements for $O_2(A^3\Sigma_u^+)(v = 0-4)$, $O_2(A^3\Sigma_u^+)(v = 2)$, and $O_2(c^1\Sigma_u^-)$ infer the values of $k_3^O/k_3^{O_2}$ ratio to be 30±30, 100±15, and 200±20, respectively (Kenner and Ogryzlo, 1980; Kenner and Ogryzlo, 1983a, 1983b; Kenner and Ogryzlo, 1984). On the other hand, Slanger et al. (1984) found a lower limit of $O_2(A^3\Sigma_u^+)(v = 8)$ quenching by O_2 must be $\geq 8 \cdot 10^{-11}$. If the results from Slanger et al. (1984) were applied to the results from Kenner and Ogryzlo (1980, 1984) for $k_3^{O_2}$, then the ratio of $k_3^O/k_3^{O_2}$would be two orders lower. This short discussion illustrates a strong scattering of this ratio. For our two potential candidates ($O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$), there is information about $k_3^O/k_3^{O_2}$ ratio for only $O_2(A'^3\Delta_u)$. Through the comprehensive analysis of known rocket experiments, López-González et al. (1992a, b, c) inferred that the upper limit of the ratio amounts to 1. Hence, our value of $k_3^O/k_3^{O_2} = 0.21^{+0.32}_{-0.12}$ agrees with this result. Consistent information from laboratory experiments on the ratio for $O_2(^5\Pi_g)$ is absent. Thus, we can propose as potential candidates for precursor both $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$; however, we are not able to identify which of these two is more preferable.

In order to illustrate the application of the newly derived fitting coefficients we compare in Figure 4 the atomic oxygen concentration from FIPEX (black line), from NRL MSISE-00 reference atmosphere model (Picone et al., 2002) (red line); calculated with McDade et al.

(1986) coefficients (blue line), and with our fitting coefficients for the two-step mechanism (green line). In the region 94-98 km, i.e. at atomic oxygen peak and volume emission peak (see Fig. 1d) fitting coefficients from this paper better then McDade coefficients (MSIS-83 case) reproduce observed values. Our fitting coefficients and fitting coefficients of McDade give similar approximation above atomic oxygen peak (~98-104 km). The shape of the calculated profiles appears slightly different, with the peak maximum at a higher altitude than the observed. In this, our result resembles the McDade results, probably because in both cases, the ratio of two reaction rates is derived, but not the rates themselves. In the lower part our results and those of McDade differ, because our C^{O2} value is larger and term with molecular oxygen dominates. Nevertheless, the atomic oxygen retrieved with our fitting coefficients satisfactorily reproduces measurements, especially at the peak.

4.3 Combined mechanism

In the most general case, the $O_2(b^1\Sigma_g^+)$ population passes through two channels: directly and via precursor. In fact, theoretical calculations from Wraight (1982) and laboratory measurements from Bates (1988) predicted a direct population with efficiencies of 0.015 and 0.03, respectively, which is not sufficient to explain the observed emissions (Bates, 1988, Greer et al., 1981; Krasnopolsky, 1986). A similar value, ε=0.02, was shown in the analysis by López-González et al. (1992b, c). We investigated a combined mechanism based on the LSF calculation and fit function (derivation in Appendix):

$$\frac{[O_2] + D_1[O]}{D_2 + \tilde{\varepsilon}(1 + D_1\,[O]/[O_2])} = \frac{A_1 k_1 [O]^2 M [O_2]}{V_{at}\left(A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]\right)}, \quad (5)$$

where, hereafter, tildes denote that these are values for combined mechanism and do not equal to the values for one-step or two-step mechanisms (Sec. 4.1 and 4.2); $D_1 = \tilde{k}_3^O/\tilde{k}_3^{O_2}$ and $D_2 = \tilde{\alpha}\tilde{\gamma}$ are the fitting coefficients, which refer to the ratio of quenching rates and $\tilde{\eta} \equiv \tilde{\alpha}\tilde{\gamma}$

total efficiency for two-step channel, respectively. The fitting coefficients were calculated for two limit cases $\tilde{\varepsilon}$=0.015 (Wraight, 1982), $\tilde{\varepsilon}$=0.03 (Bates, 1988) and for the averaged case $\tilde{\varepsilon}$=0.022.

The results for the best-fit in each case are listed in Table 3. Analogously to the two-step mechanism (Sec. 4.2), for the case of combined mechanism $\tilde{\gamma} = \tilde{\eta}/\tilde{\alpha} < 1$, hence, the precursor should satisfy $\tilde{\alpha} > 0.08^{+0.12}_{-0.04}$ (see Tab. 3). For central values of $\tilde{\alpha}$, only $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$ satisfy this criterion (see Tab. 2). At lower limit of uncertainty ($\tilde{\alpha} > 0.04$) $O_2(A'^3\Delta_u)$, $O_2(A^3\Sigma_u^+)$, and $O_2(^5\Pi_g)$ satisfy to the criterion, and considering upper limit ($\tilde{\alpha} > 0.2$), only $O_2(^5\Pi_g)$ may serve as precursor. The upper limit of the ratio $k_3^O/k_3^{O_2} < 1$ for $O_2(A'^3\Delta_u)$, derived by López-González et al. (1992a, b, c), is in agreement with our calculations ($0.231^{+0.358}_{-0.142}$). As it is noted above, the ratio for $O_2(^5\Pi_g)$ is unknown. Consequently, taking into an account both conditions, only $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$ may serve as precursor.

Figure 5 illustrates a sanity check for volume emissions derived (black lines) with fitting coefficients of McDade et al. (1986) for MSIS-83 (Fig. 5c) case and CIRA-72 case (Fig. 5d), and with our newly derived fitting coefficients for two-step (Fig. 5a) and combined ($\tilde{\varepsilon} = 0.022$) mechanisms (Fig. 5b) in comparison with measured one (red lines). All of derived volume emission profiles (black lines) were calculated based on the temperature, concentration of surrounding air, and concentration of atomic oxygen from our rocket launch. The calculations with combined mechanism (Eq. 5) and two-step energy transfer mechanism (Eq. 4) give almost identical results. The results obtained with new fitting coefficients are in satisfactory agreement with the measured volume emissions at the peak and above, whereas the McDade coefficients related to the temperature from CIRA-72 give better approximations below the volume emission peak (92 km). The coefficients of McDade related to the temperature from MSIS-83 are in better agreement with our results and are almost identical

above the volume emission peak. More independent common volume in-situ measurements are necessary to validate these results.

5. Summary and conclusions

Based on the rocket-born common volume simultaneous observations of atomic oxygen, atmospheric band emission (762 nm), and density and temperature of the background atmosphere, the mechanisms of $O_2\left(b^1\Sigma_g^+\right)$ formation were analysed. Our calculations show that one-step direct excitation is less probable by the reasons discussed above (Sec. 4.1).

For the case of the two-step mechanism, we found new coefficients for fit function in accordance with McDade et al. (1986), based on self-consistent temperature, atomic oxygen and volume emission observation. These coefficients amounted to $C^{O2}=9.8^{-5.3}_{+6.5}$ and $C^{O}=2.1^{+0.3}_{-0.6}$. C^{O2} coefficient is partially, within the error range, in agreement with C^{O2} coefficients given in McDade et al. (1986), and C^{O} coefficient is approximately one order lower. The general implication of these results is parameterisation of volume emission in terms of known atomic oxygen. This can be utilised either for atmospheric band volume emission modelling or for estimation of atomic oxygen by known volume emission. We identified two candidates for the intermediate state of O_2^*. Our results show that $O_2(A'^3\Delta_u)$ or $O_2(^5\Pi_g)$ may serve as a precursor.

Taking into account both channels of $O_2\left(b^1\Sigma_g^+\right)$ formation, we proposed a combined mechanism. In this case, atomic oxygen via volume emission or volume emission based on known atomic oxygen can be calculated by equation (5). Recommended fitting coefficients amounted to $D_1=0.231^{+0.358}_{-0.142}$ and $D_2=0.08^{+0.12}_{-0.04}$, with the efficiency of the direct channel as $\tilde{\varepsilon}=0.022$. These coefficients have a meaning of total efficiency ($\tilde{\alpha}\tilde{\gamma}$) and a ratio of quenching coefficients ($\tilde{k}_3^O/\tilde{k}_3^{O_2}$) for the two-step channel. The analysis of their values

indicates that $O_2(A'^3\Delta_u)$ and $O_2(^5\Pi_g)$ may serve as possible precursors for the two-step channel at combined mechanism. In the context of our rocket experiment, we do not have the possibility to figure out which mechanism is true. Nevertheless, we consider the combined mechanism as more relevant to nature, because it has a higher generality. This conclusion does not contradict to the current point of view that the two-step mechanism is dominant because $\tilde{\varepsilon}$ is assumed to be 1.5-3 %. Moreover, it is possible that in the reality the mechanism is much more complex and it has multi-channel or more than two-step nature. Undoubtedly, more common volume simultaneous observations of the Atmospheric Band and the atomic oxygen concentrations would be desirable to confirm and improve these results.

Appendix.

We consider photochemical equilibrium for the night-time $O_2\left(b^1\Sigma_g^+\right)$ concentration. If $O_2\left(b^1\Sigma_g^+\right)$ is produced via both channels, the equilibrium concentration is given by the following expression:

$$\left[O_2\left(b^1\Sigma_g^+\right)\right] = \frac{\tilde{\varepsilon}k_1[O]^2M + \tilde{\gamma}\tilde{k}_3^{O_2}[O_2][O_2^*]}{A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]}, \qquad (A1)$$

where the tilde denotes the combined mechanism, $A_1, k_1, k_2^{O_2}, k_2^{N_2}, k_2^{O}, \tilde{k}_3^{O_2}$ are the ratios for corresponding processes (see Tab. 1) and O_2^* is the unknown precursor.

Considering this precursor in photochemical equilibrium, we can obtain the following expression for its concentration:

$$[O_2^*] = \frac{\tilde{\alpha}k_1[O]^2M}{\tilde{A}_3 + \tilde{k}_3^{O_2}[O_2] + \tilde{k}_3^{N_2}[N_2] + \tilde{k}_3^{O}[O]}, \qquad (A2)$$

where efficiency $\tilde{\alpha}$, $\tilde{A}_3$ is the unknown spontaneous emission coefficient of O_2^* and $\tilde{k}_3^{O_2}, \tilde{k}_3^{N_2}, \tilde{k}_3^{O}$ are the unknown quenching rates for O_2^*.

Substituting A2 into A1 and into expression for volume emission we obtain:

$$V_{at} = A_1\left[O_2\left(b^1\Sigma_g^+\right)\right] =$$

$$= \frac{A_1 k_1 [O]^2 [O_2] M}{A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]} \left(\frac{\tilde{\varepsilon}}{[O_2]} + \frac{\tilde{\alpha}\tilde{\gamma}\tilde{k}_3^{O_2}}{\tilde{A}_3 + \tilde{k}_3^{O_2}[O_2] + \tilde{k}_3^{N_2}[N_2] + \tilde{k}_3^{O}[O]} \right). \qquad (A3)$$

We assume that, in analogy with two-step mechanism, a spontaneous emission $\widetilde{A_3}$ of O_2^* is much smaller than the quenching, and we utilised traditional assumption about low quenching with molecular nitrogen $\left(\tilde{k}_3^{N_2} \ll \tilde{k}_3^{O_2}\right)$, which is commonly used to analyse a potential precursor. In this case, A3 can be rearranged as follows:

$$\frac{[O_2] + \frac{\tilde{k}_3^{O}}{\tilde{k}_3^{O_2}}[O]}{\tilde{\alpha}\tilde{\gamma} + \tilde{\varepsilon}\left(1 + \frac{\tilde{k}_3^{O}}{\tilde{k}_3^{O_2}}[O]/[O_2]\right)} = \frac{A_1 k_1 [O]^2 M [O_2]}{V_{at}\left(A_2 + k_2^{O_2}[O_2] + k_2^{N_2}[N_2] + k_2^{O}[O]\right)} . \qquad (A4)$$

We defined unknown fitting coefficients $D_1 \equiv \tilde{k}_3^{O}/\tilde{k}_3^{O_2}$ and $D_2 \equiv \tilde{\alpha}\tilde{\gamma}$. Expression A4 was utilised to calculate them with LSF.

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The rocket-borne measurements and calculated data shown in this paper are available via IAP's ftp server at ftp://ftp.iap-kborn.de/data-in-publications/GrygalashvylyACP2018.

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Table 1. List of reactions with corresponding reaction rates (for three-body reactions [cm^6 $molecule^{-2}$ s^{-1}] and for two-body reactions [cm^3 $molecule^{-1}$ s^{-1}]), quenching coefficients, and spontaneous emission coefficients (s^{-1}) used in the paper.

	Reaction	Coefficient	Reference
R1	$O + O + M \xrightarrow{\varepsilon k_1} O_2(b^1\Sigma_g^+) + M$	$k_1 = 4.7 \cdot 10^{-33}(300/T)^2$ $\varepsilon - unknown$	Campbel and Gray (1973)
R2	$O_2(b^1\Sigma_g^+) + O_2 \xrightarrow{k_2^{O_2}} products$	$k_2^{O_2} = 7.4 \cdot 10^{-17} T^{0.5} e^{-\frac{1104.7}{T}}$	Zagidullin et al. (2017)
R3	$O_2(b^1\Sigma_g^+) + N_2 \xrightarrow{k_2^{N_2}} products$	$k_2^{N_2} = 8 \cdot 10^{-20} T^{1.5} e^{\frac{503}{T}}$	Zagidullin et al. (2017)
R4	$O_2(b^1\Sigma_g^+) + O \xrightarrow{k_2^O} products$	$k_2^O = 8 \cdot 10^{-14}$	Slanger and Black (1979)
R5	$O_2(b^1\Sigma_g^+) \xrightarrow{A_1} O_2 + hv(762nm)$	$A_1 = 0.0834$	Newnham and Ballard (1998)
R6	$O_2(b^1\Sigma_g^+) \xrightarrow{A_2} O_2 + hv(total)$	$A_2 = 0.088158$	Yankovsky et al. (2016)
R7	$O + O + M \xrightarrow{\alpha k_1} O_2^* + M$	$\alpha - unknown$	
R8	$O_2^* + O_2 \xrightarrow{\gamma k_3^{O_2}} O_2(b^1\Sigma_g^+) + O_2$	$\gamma - unknown$	
R9	$O_2^* + O_2, N_2, O \xrightarrow{k_3^{O_2}, k_3^{N_2}, k_3^O} prod.$	$k_3^{O_2}, k_3^{N_2}, k_3^O - unknown$	
R10	$O_2^* \xrightarrow{A_3} O_2 + hv$	$A_3 - unknown$	

Table 2. Efficiencies α of the different excited states of O_2.

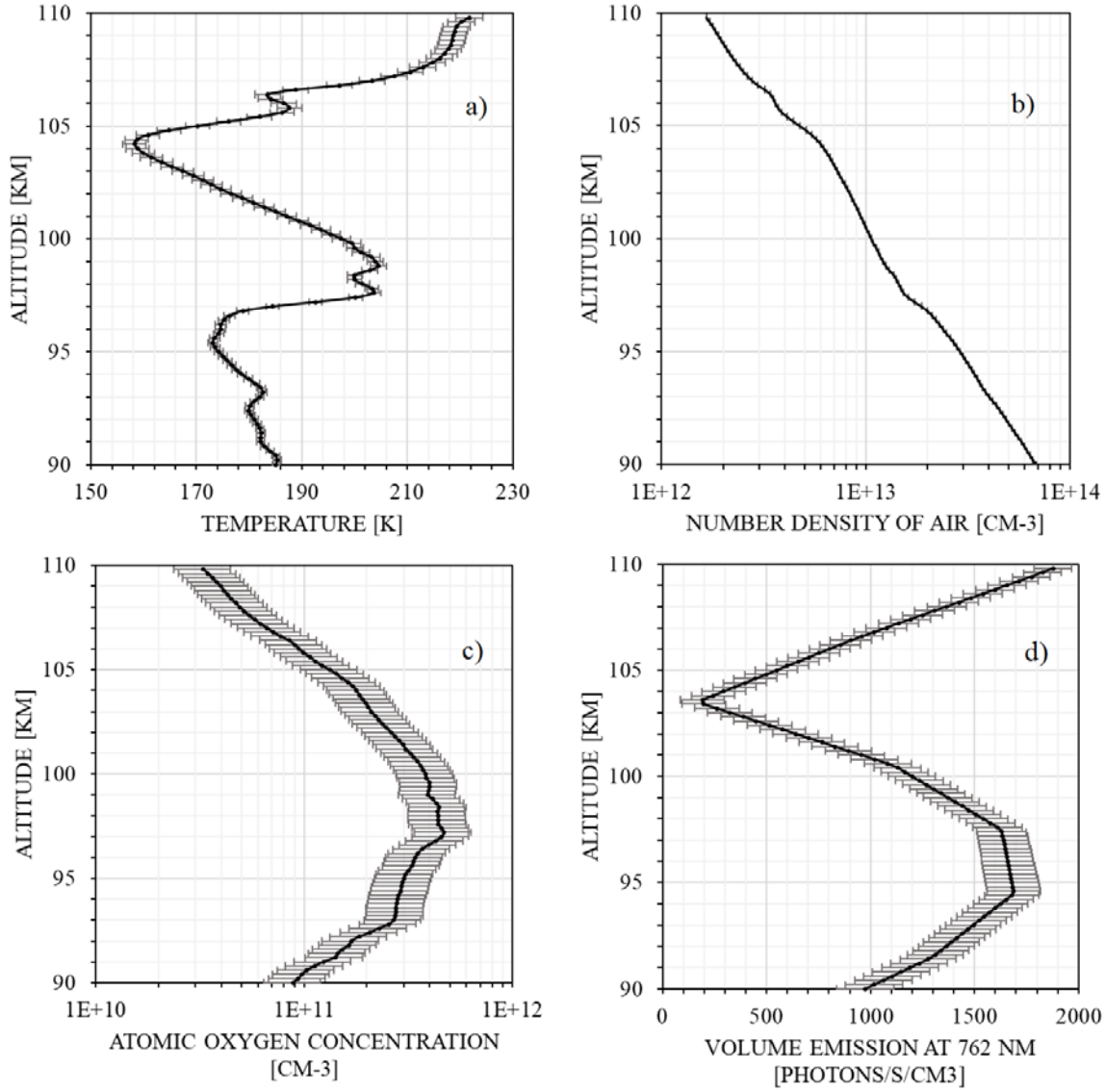
$O_2(c^1\Sigma_u^-)$	$O_2(A'^3\Delta_u)$	$O_2(A^3\Sigma_u^+)$	$O_2(^5\Pi_g)$	Reference
0.03	0.12	0.04	0.66	Wraight (1982), Smith (1984)
0.04	0.18	0.06	0.5	Bates (1988)
0.03	0.18	0.06	0.52	López-González et al. (1992a, b, c)

Table 3. Fitting coefficients for combined mechanism (Eq. 5) at different efficiencies.

	Low $\tilde{\varepsilon}$ Wraight (1982)	High $\tilde{\varepsilon}$ Bates (1988)	Averaged $\tilde{\varepsilon}$ (this work)
$\tilde{\varepsilon}$	0.015	0.03	0.022
$D_1 = \tilde{k}_3^O / \tilde{k}_3^{O_2}$	$0.211^{+0.355}_{-0.136}$	$0.397^{+0.22}_{-0.282}$	$0.231^{+0.358}_{-0.142}$
$D_2 = \tilde{\alpha}\tilde{\gamma}$	$0.087^{+0.12}_{-0.041}$	$0.073^{+0.119}_{-0.042}$	$0.08^{+0.12}_{-0.04}$

Figures.

Figure 1. Measurements of a) temperature (CONE), b) number density of air (CONE), c) atomic oxygen concentration (FIPEX), d) volume emission at 762 nm (photometer).



687 Figure 2. Quantum yield of $O_2(b^1\Sigma_g^+)$ formation ε for the case of one-step mechanism.

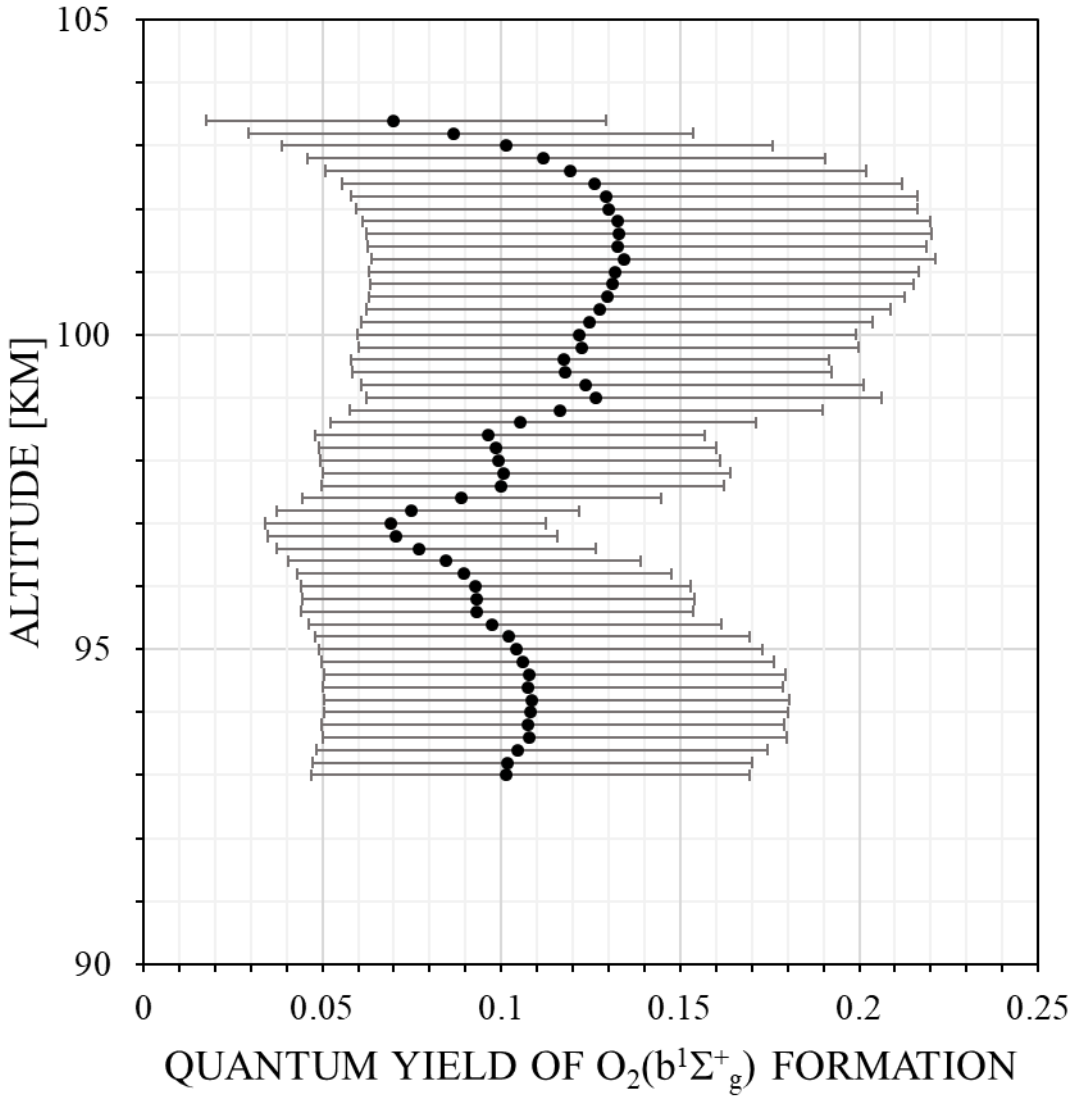


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Figure 3. RHS (dots) and least-square fit of LHS (black line) of equation (4).

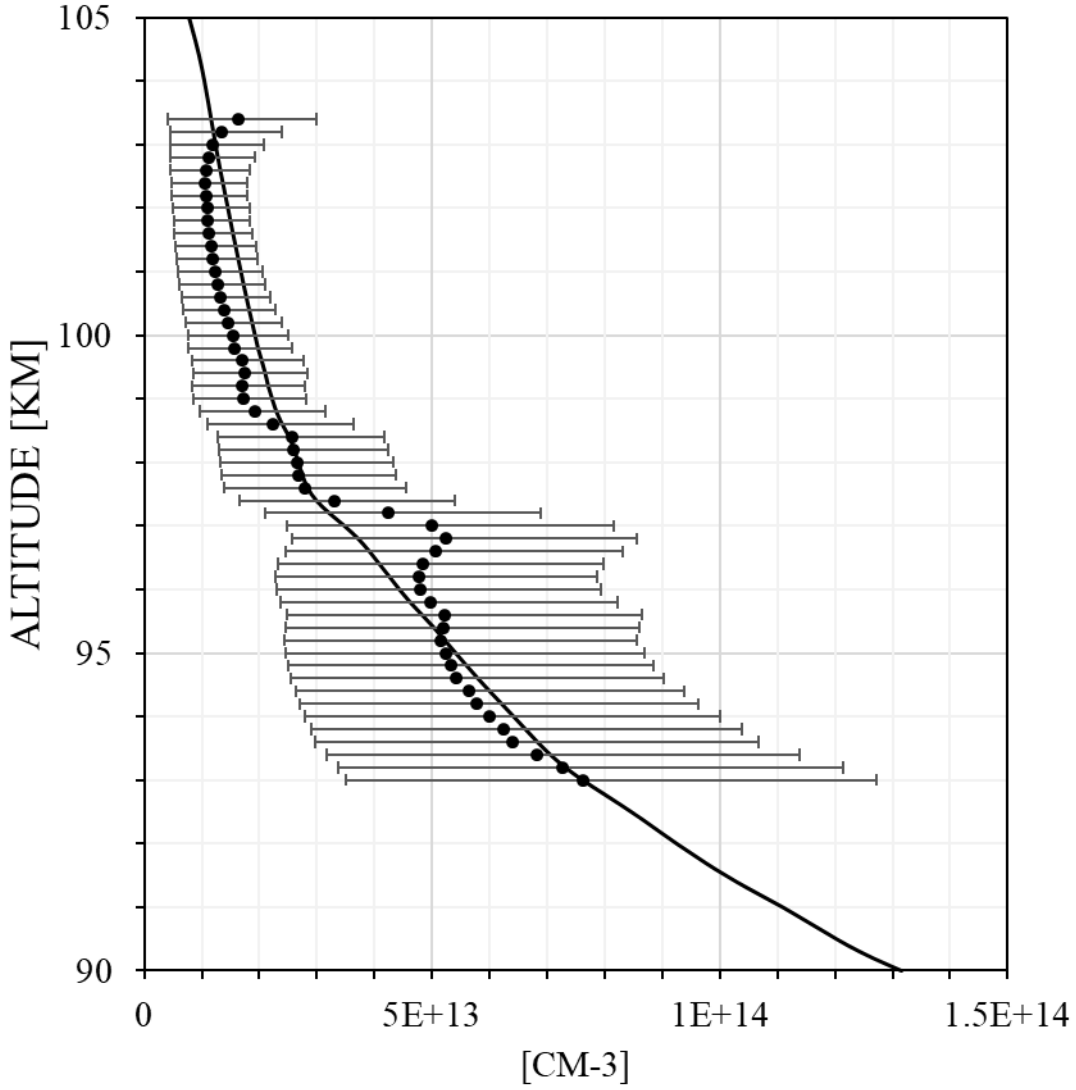

Figure 4. Atomic oxygen concentration: FIPEX (black line); model MSIS00 (red line); derived from emission observation with McDade et al. (1986) coefficients (blue line); calculated with newly derived fitting coefficients for the two-step mechanism (green line).

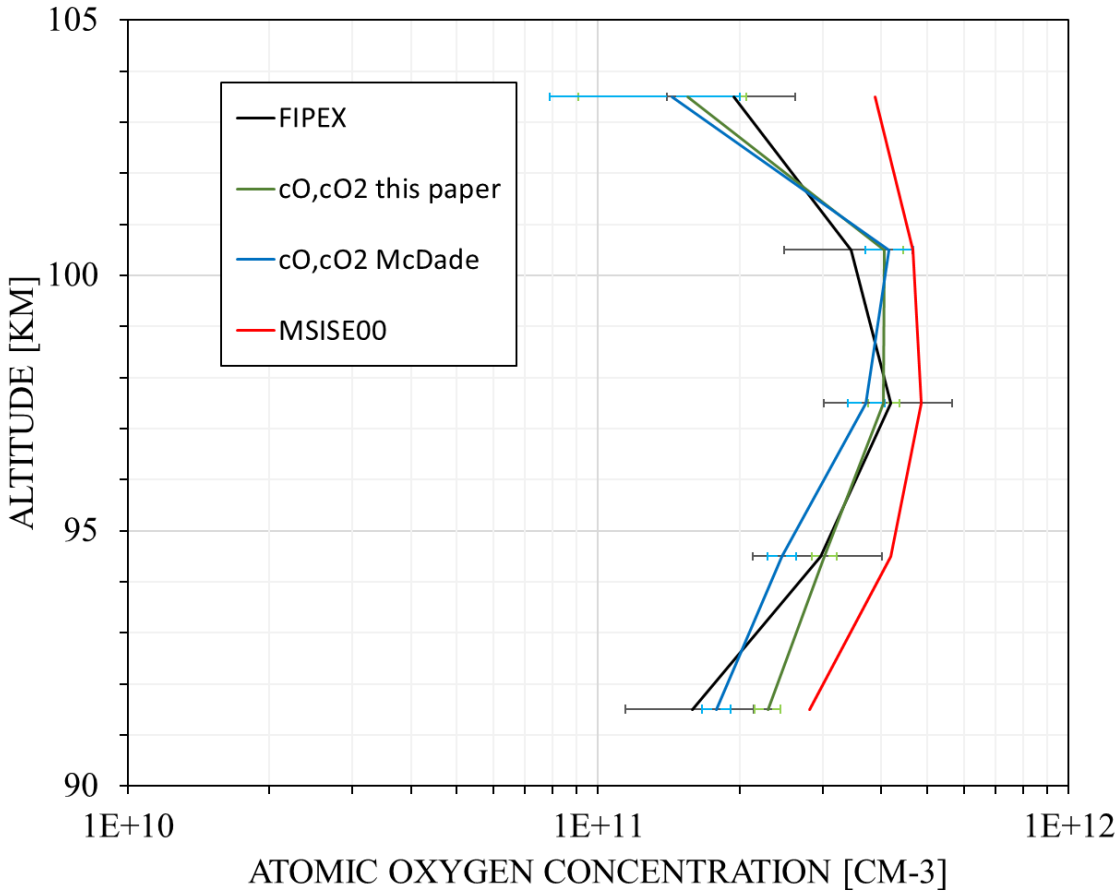

Figure 5. Volume emissions: photometer (red line); derived from atomic oxygen (black line) with a) newly derived fitting coefficients for the two-step mechanism, b) with fitting coefficients for combined mechanism, c) with McDade et al. (1986) coefficients, which correspond to the MSIS-83 temperature, and with McDade et al. (1986) coefficients, which correspond to the CIRA-72 temperatures.