The oxidation regime and SOA composition in limonene ozonolysis: Roles of different double bonds, radicals, and water

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Abstract. Volatile organic compounds (VOCs) play an important role in air quality and climate change, largely because they contribute to the formation of oxidizing compounds and secondary organic aerosol (SOA). In this study, a series of products, including peroxides and carbonyl compounds in both gaseous and particulate phases, were simultaneously detected to investigate the oxidation regime and SOA composition in limonene ozonolysis. The roles of different double bonds (DBs), radicals, and water were also examined. In our first investigation, we focused on representative oxidizing compounds produced in limonene ozonolysis, including stabilized Criegee intermediates (SCIs), OH radicals, and peroxides. The dependence of H_2O_2 and hydroxymethyl hydroperoxide (HMHP) formation on RH demonstrates that the reaction with water is an important reaction pathway for limonene SCIs, and the lower-limit SCI yields of endocyclic and exocyclic DBs were estimated to be ~0.24 and ~0.43, respectively. The OH yield was determined by adding sufficient amounts of an OH scavenger, and the OH yields of endocyclic and exocyclic DBs were ~0.65 and ~0.24, respectively. These results indicate that in limonene ozonolysis the endocyclic DB is inclined to generate OH radicals through hydroperoxide channel, while the exocyclic DB has a higher fraction of forming SCIs. Besides, other gas-phase and particle-phase peroxides were also studied in this work. The formation of peroxyformic acid (PFA) and peroxyacetic acid (PAA) was promoted significantly by increasing RH and the degree of oxidation, and the discrepancy between

the experimental and model results suggested some missing formation pathways. Considerable generation of H_2O_2 from SOA in the aqueous phase was observed, especially at high $[O_3]/[limonene]$, which was mainly attributed to the hydration and decomposition of unstable peroxides in SOA such as peroxycarboxylic acids and peroxyhemiacetals. Different DBs and OH scavengers had a large impact on the particulate peroxides, and their stability indicated that the types of peroxides in SOA changed under different conditions. As for the contribution of peroxides to SOA, the results demonstrated that the mass fraction of particulate peroxides in limonene SOA was less than 0.2 at low $[O_3]/[limonene]$, while the mass fraction was 0.4–0.6 at high $[O_3]/[limonene]$. The partitioning behavior of peroxides showed that multi-generation oxidation helped produce more low-volatility peroxides, which partially explained the higher SOA yield. The partitioning behavior of carbonyls was also examined and the experimental partitioning coefficients (K_p) were found to be typically several orders of magnitude higher than the theoretical values. This study provided new insights into the oxidation regime and SOA composition in limonene ozonolysis, and limonene showed its specificity in many aspects when both endocyclic and exocyclic DBs were ozonated. We suggest that the atmospheric implications of terpenes containing more than one DB and the SOA composition, especially particulate peroxides, need further study.

1 Introduction

As an important monoterpene, limonene has a high emission rate from both biogenic and anthropogenic sources that is second only to pinene (Atkinson and Arey, 2003; Clausen et al., 2001; Fellin and Otson, 1994; Griffin et al., 1999; Guenther et al., 1995; Lamb et al., 1993; Seifert et al., 1989; Sindelarova et al., 2014; Wolkoff et al., 2000). Total monoterpene emissions are estimated to be 50 Tg C yr^{-1}, and limonene might comprise about 20% of that (Stroud et al., 2005). In addition to massive emissions from vegetation, the extensive utilization of limonene in household and industrial processes makes it non-negligible in the atmosphere. Compared with its

isomers α-pinene and β-pinene, an obvious structural feature of limonene is that it has two different double bonds (DBs), an endocyclic DB and an exocyclic DB, which complicates the behavior and the fate of limonene in the atmosphere. Ozonolysis is an important reaction pathway of limonene and it serves as a source of free radicals, intermediate products, and aerosol. The first step in alkene ozonolysis is the addition of O_3 to the carbon-carbon double bond, forming an energy-rich primary ozonide (POZ). The POZ decomposes into two sets of carbonyls plus carbonyl oxides, which are called excited Criegee intermediates (ECIs) (Criegee, 1975; Fenske et al., 2000; Gutbrod et al., 1996, 1997; Kroll et al., 2001a, b). ECIs can isomerize through the hydroperoxide channel followed by OH production, rearrange to esters with subsequent decomposition, or undergo collisional stabilization forming stabilized Criegee intermediates (SCIs) (Cremer et al., 1993; Presto and Donahue, 2004; Richard et al., 1999; Wegener et al., 2007).

In the last few years, the oxidizing compounds formed in alkene ozonolysis have received much attention because they are able to remove a series of trace gases and contribute to atmospheric oxidation capacity (Mӧller, 2009; Prinn, 2003; Taatjes et al., 2013). Representatives of these compounds include OH radicals, SCIs, peroxides, etc. Investigations into the oxidation regime of alkene ozonolysis are essential to understanding the formation of these critical products and evaluating the actual impact of this reaction in the atmosphere. As an important source of OH radicals needed for OH-initiated reactions occurring in darkness (Kroll et al., 2001a, b), OH formation pathways in alkene ozonolysis have been extensively studied and the major pathway is considered to be the unimolecular decomposition of ECIs. Furthermore, some studies suggest that SCIs could also generate OH radicals through self-decomposition or reactions with other species (Anglada et al., 2002; Hasson et al., 2003; Kroll et al. 2001a, b; Tillmann et al., 2010; Zhang and Zhang, 2005). SCIs generated during the process have been shown to have a lifetime, usually on the order of sub-seconds (Kroll et al., 2001a, b; Long et al., 2018; Mauldin et al., 2012), that is sufficiently long to react with other trace species, such as H_2O, SO_2,

65   NO_2, carbonyls, alcohols, and carboxylic acids. (Sakamoto et al., 2017; Sipilä et al. 2014; Yao et al., 2014). A

number of studies have focused on the reactions of SCIs containing 3 or fewer carbon atoms, and their

unimolecular decomposition and reactions with water show strong structure dependence. The reactions with

water monomer and water dimer are the main reaction pathways for CH_2OO and *anti*-CH_3CHOO, while the

unimolecular decomposition would be more important for *syn*-CH_3CHOO and $(CH_3)_2COO$ (Huang et al., 2015;

70   Lin et al., 2016; Welz et al., 2012, 2014; Taatjes et al., 2013). However, the reaction mechanisms for larger SCIs

formed from biogenic alkenes ozonolysis are still unclear, since it is difficult to synthesis and observe them

directly. Furthermore, reactions of alkenes with ozone can also generate a considerable amount of peroxides,

which are considered to be important due to their oxidizability and role as a radicals reservoir. H_2O_2 is the most

crucial oxidant to oxidize S (IV), forming sulfuric acid and sulfate in the aqueous phase (Calvert et al., 1985;

75   Penkett et al., 1979; Peña et al., 2001). And as a species accounting for 40%−50% of total global organic

peroxides (Crounse et al., 2006; Khan et al., 2015), peroxycarboxylic acids (RC(O)OOH) play an important role

in promoting atmospheric oxidation capacity and enhancing the acidity of the aqueous phase.

One important reason why limonene chemistry has attracted much attention is its high SOA formation potential,

which is higher than that of α-pinene since limonene is doubly unsaturated (Andersson-Sköld and Simpson 2001;

80   Kroll and Seinfeld, 2008; Lane et al., 2008; Lee et al., 2006). Although progress has been made in simulating

SOA formation with the theory of gas-particle partitioning, there are still large discrepancies between model and

experimental results (Cocker et al., 2001; Griffin et al., 1999; Hoffmann et al., 1997; Odum et al., 1996; Pankow,

1994; Presto et al., 2005; Pye and Seinfeld, 2010). Laboratory studies of SOA formation in limonene ozonolysis

mainly focused on the aerosol yields under different conditions and the identification of products in the particle

85   phase (Calogirou et al., 1999; Glasius et al., 2000; Grosjean et al., 1992, 1993; Leungsakul et al., 2005; Ng et al.,

2006; Pathak et al., 2012), however, the composition of limonene SOA still needs detailed study, especially the

effect of different DBs. As a double-unsaturated terpene, the SOA formation process of limonene could be more complicated than that for single-unsaturated terpene, since multi-generation oxidation has significant influence on SOA. In this study, two classes of species, peroxides and carbonyls are studied for their behaviors in limonene SOA formation. Over the past few years, analyses of organic peroxides suggest that they are an important component in aerosol (Docherty et al., 2005; Heaton et al., 2007; Li et al., 2016; Mertes et al., 2012; Pathak et al., 2012), and particulate peroxides could cause negative health effects when they penetrate into lungs (Chen et al., 2011; Fischer and Maier, 2015; Gupta et al., 2014; Verma et al., 2009; Wragg et al., 2016). The reactive uptake and particle-phase reactions of carbonyls are believed to be responsible for fractions of aqueous SOA formation (Ervens et al., 2011; Mcneill et al., 2012), especially for the dicarbonyls glyoxal and methylglyoxal. Until now, there has been little research into the contribution of peroxides and carbonyls to limonene SOA and their behaviors in aerosol formation are only vaguely understood, to the best of our knowledge.

The focus of this study is to investigate the oxidation regime and SOA composition in limonene ozonolysis, especially for the roles of different DBs, radicals, and water. We consider the generation of SCIs, OH radicals, and peroxides in discussions of the oxidation regime in the reaction system. Furthermore, peroxides and carbonyls are taken as representatives to study their behaviors in SOA formation.

2 Experimental

2.1 Chemicals

This study used the following: R-(+)-Limonene (Sigma-Aldrich, ≥ 99.0%), 2-butanol (Sigma-Aldrich, 99.5%), cyclohexane (Alfa Aesar, ≥ 99.9%), potassium iodide (KI, Alfa Aesar, 99.9%), hydrogen peroxide (H_2O_2, Alfa

Aesar, 35wt.%), ortho-phosphoric acid (H$_3$PO$_4$, Fluka, 85–90%), hemin (Sigma, ⩾ 98.0%), 4-hydroxyphenylacetic acid (Alfa Aesar, 99%), formaldehyde (Sigma-Aldrich, 37wt.%), acetaldehyde (Amethyst Chemicals, 40wt.%), acetone (Fluka, ⩾ 99.7%), hydroxyacetone (Sigma-Aldrich, 90%), glyoxal (Sigma-Aldrich, 40wt.%), methylglyoxal (Sigma-Aldrich, 40wt.%), 2-butanone (Alfa Aesar, ⩾ 99%), acetonitrile (Alfa Aesar, ⩾ 99.7%), 2,4-dinitrophenyl hydrazinel (DNPH, TCI, 50wt.%), ammonia solution (NH$_3$ H$_2$O, Beijing Tongguang Fine Chemicals Company, 25.0–28.0%), ammonium chloride (NH$_4$Cl, Beijing Tongguang Fine Chemicals Company, ⩾ 99.5%), sulfuric acid (H$_2$SO$_4$, Xilong Chemical Company, 95.0–98.0%), ultrapure water (18MΩ, Millipore), N$_2$ (⩾ 99.999%, Beijing Haikeyuanchuang Practical Gas Company Limited, Beijing, China), O$_2$ (⩾ 99.999%, Beijing Haikeyuanchuang Practical Gas Company Limited, Beijing, China), polytetrafluoroethylene (PTFE) filter membrane (Whatman Inc., 47 mm in diameter), and quartz microfiber filters (Whatman Inc.).

2.2 Apparatus and procedures

A flow tube reactor (2 m length, 70 mm inner diameter, quartz wall), equipped with a water jacket for controlling the temperature, was used to investigate the ozonolysis of limonene. All the experiments were conducted at 298±0.5 K in darkness. O$_3$ was generated by O$_2$ photolysis in a 2 L quartz tube with a low-pressure Hg lamp, and the detailed quantification method of O$_3$ was described in our previous study (Chen et al., 2008). H$_2$O$_2$ produced by UV irradiation of O$_2$ and trace water was measured in control experiments and deducted from the results. Limonene gas was generated by passing N$_2$ flow over liquid limonene in a diffusion tube at the selected temperature, and OH scavenger (2-butanol or cyclohexane) gas was generated with a bubbler. The concentrations of limonene and the OH scavenger were determined by gas chromatography with a flame ionization detector (GC-FID, Agilent 7890A, USA). Water vapor was generated by passing N$_2$ through a water

bubbler, which contained a carborundum disc submerged in ultrapure water (18 MΩ). The mixing gas, including limonene, OH scavenger, ozone, and dry or wet synthetic air (80% N_2 and 20% O_2), was successively introduced into the reactor. With a total flow rate of 2 standard L min^{-1}, the residence time was estimated to be 240 s.

To explore the reaction mechanism of endocyclic and exocyclic DBs ozonolysis and the effect of multi-generation oxidation in limonene ozonolysis, we conducted two sets of experiments with different ratios of ozone to limonene concentration. In the following, $[O_3]$ denotes the concentration of ozone, [limonene] denotes the concentration of limonene, and $[O_3]$/[limonene] denotes the concentration ratio of ozone to limonene. In the low $[O_3]$/[limonene] experiments, the initial concentrations of limonene and ozone were ~280 ppbv and ~500 ppbv, respectively, whereas in the high $[O_3]$/[limonene] experiments, the initial concentrations of limonene and ozone were ~183 ppbv and ~19 ppmv, respectively. In both sets of experiments, sufficient quantities of 2-butanol and cyclohexane were added to scavenge OH radicals in the RH range of 0%–90%. In the tables and figures, the low-ratio and high-ratio experiments are denoted by L and H, respectively, and the conditions in the absence of scavenger, in the presence of 2-butanol, and in the presence of cyclohexane are denoted by No-sca, 2-But, and C-hex, respectively. The experimental conditions are listed in Table 1.

According to previous studies, the rate constant of the endocyclic DB reaction with ozone in limonene ozonolysis is 2×10^{-16} cm^3 molecule^{-1} s^{-1} (Atkinson, 1990; Shu and Atkinson, 1994), while the exocyclic DB reaction with ozone is about 30 times slower than the endocyclic DB reaction (Zhang et al., 2006). Based on those rate constants, we estimate that at low $[O_3]$/[limonene], less than 1% exocyclic DB is ozonated, so this situation mainly represents first-generation oxidation. In this circumstance, because the ozone concentration is low, the OH reaction would impact the amount of limonene consumed by O_3. In the presence of OH scavenger, ~42% endocyclic DB reacted with O_3, while in the absence of scavenger, ~38% endocyclic DB reacted with O_3.

150   At high [O$_3$]/[limonene], more than 99% endocyclic DB and about 51% exocyclic DB reacted with ozone, and

since the ozone concentration in this situation was high, the OH effect on the ozonolysis was presumed to be

unimportant. The latter condition, which contained multi-generation oxidation processes, was more likely to

occur since the ratio of [O$_3$] to [limonene] was similar to the ratio in the real atmosphere.

It should be noted that one advantage of flow tube reactor is that the wall would be in equilibrium with the gas

155   phase after a stationary period. According to our observations, this process usually requires about 2 h. Thus, to

stabilize the system and diminish the wall effect to the extent possible, we ran the reactor for 2 h prior to taking

measurements. And after the experiments, the reactor was rinsed with ultrapure water and blown dry with N$_2$.

2.3 Products analysis

To better investigate the gas-particle partitioning of products formed in limonene ozonolysis, we analyzed the

160   gas-phase and particle-phase products simultaneously. We measured total peroxides and a series of

low-molecule-weight (LMW) peroxides. In the following discussion we regard those peroxides that could be

detected by high performance liquid chromatography (HPLC) to be LMW peroxides, while took the peroxides

undetermined by HPLC to be high-molecule-weight (HMW) peroxides. For particle-phase peroxides detection,

a PTFE filter was used for the SOA collection, and the mass of SOA was measured by semi-micro balance

165   (Sartorius, Germany). Since the control experiment showed that a long-duration collection led to the loss of

some peroxides in particles, the collection time was controlled to be 3 h for each filter, and the accuracy of the

particulate products analysis was discussed in the Supplement. Each loaded PTFE filter was extracted with 20

mL H$_3$PO$_4$ solution (pH 3.5) using a shaker (Shanghai Zhicheng ZWY 103D, China) at 180 rpm and 4 ℃ for 15

min, and the extraction efficiency was confirmed in our previous work (Li et al., 2016). For gas-phase peroxides

170   detection, the air samples through the filter were collected in a glass coil collector at a temperature of 4 ℃ with

H$_3$PO$_4$ solution (pH 3.5) serving as the rinsing solution. The coil collector is around 30 cm long and its effective length is about 100 cm. The coil is similar with that used in earlier studies (Grossmann et al., 2003; Sauer et al., 1996, 1997), and the diagram of the thermostatic coil collector is provided in the Supplement.

The method for detecting the peroxides was reported in our previous studies (Hua et al., 2008; Li et al., 2016)and is only briefly described here. LMW peroxides were analyzed by HPLC (Agilent 1100, USA) coupled with post-column derivatization and fluorescence detection on line. Peroxides separated by column chromatography reacted with *p*-hydroxyphenylacetic acid (POPHA) under the catalysis of hemin forming POPHA dimers, and then the dimers were quantified by fluorescence detector. The standard solution of peroxides was prepared and used for calibration in each measurement. The detection limit for the gas-phase LMW peroxides was about 22 pptv, and in the particle phase the detection limit was around 0.068 ng/μg for H$_2$O$_2$. The accuracy of this method was estimated to be around 7% with the precision of the measurement results usually within 20%. The concentration of total peroxides (H$_2$O$_2$, ROOH, and ROOR$'$) was determined by the iodometric spectrophotometric method, which is based on the reaction of peroxides and iodide ions (Docherty et al., 2005; Mutzel et al., 2013). Briefly, excessive KI solution was added to samples purged of O$_2$.After remaining in darkness 12–24 h for the derivatization, the produced I$_3^-$ ions were quantified at 420 nm by UV/VIS spectrophotometer (SHIMADZU UV-1800, Japan). The efficiency of the total peroxides measurement was discussed in our previous work (Li et al., 2016) and the detection limit of the total peroxides was about 0.92 ppbv in the gas phase and 0.025 μg/μg in the particle phase. The accuracy of this method was estimated to be 10% with the precision of the measurement results usually within 20%.

To measure the particle-phase carbonyls, SOA was collected onto a quartz microfiber filter for 3 h. After collection, the filter was placed upside down in a conical flask, and then 5 mL acetonitrile, 1 mL DNPH saturated solution, and 50 μL H$_2$SO$_4$ solution (0.25 M) were added into the flask in sequence. The flask was

shaken at 180 rpm and 4 ℃ for 3 h and kept in darkness for 12–24 h prior to detection. For the gas-phase carbonyls measurement, gas through the filter was directly introduced into a Horibe tube, which was placed in a cold trap (Beijing Tiandijingyi TH-95-15-G, China) at about -98 ℃ to freeze the products in the tube. The Horibe tube was composed of an inlet tube (25 cm length, 4 cm O.D.), a coil (7 laps, 1 cm O.D.), and an outlet tube with a carborundum disc. The diagram of the Horibe tube is shown in the Supplement. After collection, 10 mL acetonitrile was added to rinse the inside of the Horibe tube to dissolve carbonyl compounds, and then the solution was mixed with DNPH saturated solution and H_2SO_4 solution to derivatize for 12–24 h in darkness. The derived samples were analyzed by HPLC with UV detection (Agilent 1100, USA). Details of this process are provided in our previous report (Wang et al., 2009). The detection limit of carbonyl compounds was about 0.3 ppbv in the gas phase, and in the particle phase it was 0.15 ng/μg in terms of formaldehyde (FA). The accuracy of this method was estimated to be 8% and the precision of the measurement results was within 30%. Since the yields of both acetone (ACE) and methylglyoxal (MGL) were low, especially under some conditions, their uncertainty might be high and the precision was within 80%.

2.4 Wall loss experiments

To increase the accuracy of the results, we designed and conducted a series of control experiments to quantify the wall loss effect, including gaseous products and aerosol wall loss evaluation. Gas containing peroxide constituents was generated by passing N_2 through a diffusion tube, at a controlled temperature of 4 ℃, which contained certain peroxide solution. The synthetic method of multiple organic peroxides was described in our previous work (Huang et al., 2013). Gas containing carbonyl constituents was prepared by injecting liquid substance into an evacuated steel canister (15 L, Entech Instrument). Then N_2 was added constantly until the pressure in the canister reached 30 psi. The outlet of the canister was linked with a mass flow controller that

regulated the gas flow rate. The gas containing peroxides or carbonyls was mixed with synthetic air at different

215   RH levels and introduced into the reactor at a rate of 2 standard L min^{-1}, and the concentrations of peroxides and

carbonyls were controlled to be similar to the concentrations of products observed in limonene ozonolysis. After

the gas mixture was introduced into the flow tube, a period of approximately 2 h was needed for the gas and

wall to become balanced, and then the measurements were begun. The aerosol wall loss experiment used

two-stage reaction equipment that contained two flow tube reactors. In order to explore the wall loss effect on

220   pure SOA, the first reactor was used to generate aerosol where limonene and O_3 had sufficient time to react

completely. The gas containing aerosol out of the first reactor was mixed with synthetic air at different RH

levels and introduced into the second reactor with a rate of 2 standard L min^{-1}. The particles at the inlet and

outlet of the second reactor were collected on PTFE filters and measured by balance to calculate the SOA

concentration. The wall loss fraction of the gas-phase constituent or SOA was determined as the difference

225   between the inlet and outlet concentrations divided by the inlet concentration as Eq. (S1) (in the Supplement).

The wall loss experiments were conducted in the RH range of 0%–90%, and the profiles of loss fractions as a

function of RH could be used to correct products and SOA yields to diminish the wall loss effect.

3 Results and discussion

3.1 Wall loss correction

230   The wall loss fractions of LMW peroxides increased with increasing RH, and this tendency was especially

obvious for peroxyformic acid (PFA) and peroxyacetic acid (PAA), whose loss fractions increased successively

with RH. For H_2O_2 and hydroxymethyl hydroperoxide (HMHP), their wall loss fractions went up quickly above

50% RH, yet they did not have a large change below 50% RH. Generally speaking, HMHP had the highest wall

loss fraction, reaching ~0.25 at 90% RH, while PFA had the lowest wall loss fraction. The average of these

235   LMW peroxides loss fractions was used to correct the wall loss effect for gas-phase HMW peroxides.

As for carbonyl compounds, the variation of their wall loss fractions with RH was not very obvious. The loss

curves of FA, acetaldehyde (AA), and ACE were somewhat irregular. The wall loss fractions of hydroxyacetone

(HACE), glyoxal (GL), and MGL were lowest at 10% RH, and then they gradually arose with increasing water

vapor concentration. The higher loss fractions of GL and MGL compared with other carbonyls could be

240   attributed to their inclination towards hydration. FA had the lowest wall loss fraction around 0.03, while MGL

had the highest loss fraction around 0.12.

The wall loss effect on SOA was also investigated from 0% to 90% RH. The SOA wall loss fraction was ~0.06

at the dry condition, then it increased slightly with water vapor concentration. The SOA loss fraction was ~0.11

and ~0.17 at 50% RH and 90% RH, respectively.. The dependencies of the peroxides, carbonyls, and SOA wall

245   loss fractions on RH are provided in the Supplement.

3.2 SCIs generation

The reaction between SCIs and water, which produces α-hydroxyalkyl hydroperoxides (HAHPs) decomposing

to H_2O_2, carbonyls, and carboxylic acids, is thought to be an essential source of these compounds and serves as

a principle source of H_2O_2 formation without light (Becker et al., 1990, 1993; Gäb et al., 1985; Sauer et al.,

250   1999). Figure 1 shows the dependence of H_2O_2 yield and HMHP yield on RH, and the six profiles in each

subgraph represent conditions at low or high $[O_3]/[limonene]$ in the presence or absence of the OH scavenger

(2-butanol or cyclohexane). The molar yield used here is defined as the ratio of the product's number of moles

to the number of moles of limonene consumed, as given by Eq. (S2) (in the Supplement). It was obvious that,

even in different cases, the variations of H_2O_2 yield and HMHP yield with RH had similar tendencies, both of

255   which increased significantly from 0% RH to 70% RH, and then they approached the limiting values. There was

little effect of OH scavenger. At low $[O_3]/[\text{limonene}]$, the maximum H_2O_2 yield was ~24.00% without OH

scavenger, ~24.60% with 2-butanol, and ~22.95% with cyclohexane. At high $[O_3]/[\text{limonene}]$, the maximum

H_2O_2 yield reached ~41.20% without OH scavenger, ~41.80% with 2-butanol, and ~40.50% with cyclohexane.

As for HMHP, its yield was much higher at high $[O_3]/[\text{limonene}]$ (~ 5.43%) than at low $[O_3]/[\text{limonene}]$ (~

260   0.62%), with the specific information found in the following picture. It is usually believed that H_2O_2 has two

generation pathways: HO_2 self-reaction and HAHPs decomposition. The former pathway is considered as a

main contributor to H_2O_2 during daytime, while the latter is regarded to occur without photochemistry. In the

reaction system discussed here, when we applied a box model coupled with the limonene reaction mechanism

extracted from Master Chemical Mechanism (MCM) v3.3 (http://mcm.leeds.ac.uk/MCMv3.3.1) to simulate the

265   reaction, the yield of H_2O_2 formed from the HO_2 self-reaction was estimated to be less than 0.1% under both dry

and wet conditions. When the 2-butanol or cyclohexane reaction mechanism was taken into consideration, the

contribution of this pathway to H_2O_2 formation was still very limited, hence, it was assumed that the HO_2

self-reaction was not important for H_2O_2 generation in this reaction system. Under dry condition, a small

amount of water vapor might desorb from the flow tube wall and participated in reactions resulting in a small

270   amount of H_2O_2 formation. Thus, the high H_2O_2 yield observed here might be mainly attributed to SCIs reaction

with water.

From 0% to 70% RH, both the H_2O_2 yield and HMHP yield were promoted significantly by increasing RH,

indicating that the reaction with water gradually became be dominant for limonene SCIs. Jiang et al. (2013)

suggested that the formation of HAHPs was the most favorable pathway for limonene SCIs reaction with H_2O,

275   and the subsequent decomposition of HAHPs was thought to be prior to generating aldehyde and H_2O_2 (Chen,

2016; Kumar, 2014). Although theoretical calculation results indicate that HAHPs decomposition is slow, some

studies proved that water and acid molecules could greatly promote the decomposition process (Anglada et al., 2002; Anglada et al., 2011; Aplincourt and Anglada, 2003; Crehuet et al., 2001), and H_2O_2 formation from HAHPs decomposition occurs very rapidly (Chen et al., 2016; Winterhalter et al., 2000). The fact that few HAHPs larger than HMHP were identified in alkene ozonolysis also provided evidence for the rapid decomposition of large HAHPs. Above 70% RH, the appearance of the limiting values of H_2O_2 yield and HMHP yield suggested that the water vapor concentration was high enough that the bimolecular reaction of SCIs and H_2O suppressed the unimolecular decomposition of SCIs and the reactions of SCIs with other products. In previous studies, the unimolecular decomposition of SCIs and the reactions of SCIs with water showed strong structure dependence. For CH_2OO and $anti$-CH_3CHOO, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were less than 1 ms, while their lifetimes of unimolecular reaction were much longer (Lin et al., 2016; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2012). For syn-CH_3CHOO and $(CH_3)_2COO$, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were more than 100 ms, while their lifetimes of unimolecular reaction were just a few milliseconds at 298 K (Drozd et al., 2017; Huang et al., 2015; Long et al., 2018; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2014). For larger SCIs their atmospheric lifetimes of unimolecular reaction and bimolecular reactions with H_2O and $(H_2O)_2$ were not quantified. Tillmann et al. (2010) inferred that about 46% of ECIs formed from α-pinene ozonolysis would stabilize and at 44% RH about 65% of SCIs formed H_2O_2 after reacting with water and 28% of SCIs decomposed. Yao et al. (2014) showed that a large fraction of ECIs formed from α-cedrene were stabilized and the unimolecular decomposition of SCIs was suppressed by their bimolecular reactions. Although we could not estimate the rates of SCIs decomposition and their reaction with water, the results here prove that the reaction with water was an essential route for limonene SCIs, and the rapid decomposition of HAHPs made an important contribution to H_2O_2 formation.

According to the experimental conditions elaborated in Sect. 2.2, we tried to calculate the contribution of

300   endocyclic DB and exocyclic DB ozonolysis to H_2O_2 and HMHP formation, and furthermore, infer the SCIs

generated in the ozonolysis of different DBs. The predicted SCI yield was derived by combining the limiting

yields of H_2O_2 and HMHP together, as in Eq. (S3) (in the Supplement), based on the assumption that the

limiting yield of H_2O_2 was equal to the large SCI yield (Hasson, 2001a, b). Considering that a portion of SCIs

might undergo unimolecular decomposition a lower-limit yield of SCIs formed in limonene ozonolysis was

305   estimated here. Additionally, as described in Sect. 3.3, the results demonstrated that the OH yield was not

obviously affected by RH, and we speculated that the fraction of SCIs that decomposed and formed OH radicals

was small. It was observed that the OH scavenger did not have a large impact on the SCIs measurement results,

while big difference existed between the ozonolysis of two DBs. The SCI yield of the endocyclic DB ozonolysis

was around 24.45%, yet the exocyclic DB ozonolysis had a larger stabilization fraction of ECIs, which was

310   about 42.90%. It should be noted that, since our experiments just isolated the ozonolysis of different DBs

according to their discrepancy in reaction rates, the SCI yield of the exocyclic DB ozonolysis estimated here

was not for a specific product formed from endocyclic DB ozonolysis but was the sum of first-generation

products with a remaining double bond, as shown in Fig. 2. The results meant that, even though exocyclic DB

ozonolysis is much slower than endocyclic DB, it plays a non-negligible role in generating SCIs in limonene

315   ozonolysis.

3.3 OH radicals generation

In this study, OH yield was determined by adding sufficient 2-butanol as a scavenger and measuring how much

2-butanone was generated. Aschmann et al. (2002) suggested that using 2-butanol to measure OH formation

would be more accurate than using cyclohexane. The yield of 2-butanone formed from the reaction of 2-butanol

320   and OH was detected previously (Aschmann et al., 2002; Baxley and Wells, 1998; Chew and Atkinson, 1996),

and we used 0.66, which is the average of the reported values (Forester and Wells, 2011), to calculate OH

radicals yield. The results showed that the yield of OH radicals produced from endocyclic DB ozonolysis was

0.65 ± 0.21, while for exocyclic DB, the OH yield was 0.24 ± 0.13. The OH generation did not show a clear

dependence on water vapor concentration. The OH yield of endocyclic DB ozonolysis was within the range of

325   published values, while the OH yield of exocyclic DB ozonolysis was determined to be higher than that reported

by Herrmann et al. (2010). The effect of water on OH formation in alkene ozonolysis is still under debate as

there are discrepancies among the existing publications. Anglada et al. (2002) used quantum mechanical

calculations to indicate that water could increase OH production, and Tillmann et al. (2010) reported higher OH

yield under humid condition. Nevertheless, more studies showed that OH formation was independent of water

330   vapor concentration (Aschmann et al., 2002; Atkinson and Aschmann, 1993; Atkinson et al., 1992; Berndt et al.,

2003; Forester and Wells, 2011; Hasson et al., 2003). The fact that the OH yields of both unsaturated bonds were

not obviously affected by RH suggests that the major pathway of OH generation in limonene ozonolysis is the

decomposition of ECIs through hydroperoxide channel. The possibility that other OH formation pathways also

exist cannot be totally excluded, but we speculate that any other pathways are not significant. Hence, it could be

335   concluded that in limonene ozonolysis, the endocyclic DB is inclined to generate OH radicals through ECIs

decomposition, while the exocyclic DB has a higher fraction of stabilization forming SCIs.

3.4 Peroxycarboxylic acids generation

In our experiments, two kinds of peroxycarboxylic acids PFA ($CH(O)OOH$) and PAA ($CH_3C(O)OOH$) were

observed, and the yields of PFA and PAA both showed large discrepancies between the low-ratio and high-ratio

340   experiments (Fig. 3). The yields of PFA and PAA at high $[O_3]/[limonene]$ were about three times as much as

those at low [O_3]/[limonene]. When the exocyclic DB was ozonated, the generation of PFA and PAA was enhanced to a large extent compared with only endocyclic DB ozonolysis, indicating that exocyclic DB ozonolysis had an important impact on PFA and PAA. As far as we know, this was the first time that such high yields of PFA and PAA in alkene ozonolysis have been reported. For both PFA and PAA, the highest molar

345   yields were observed when no OH scavenger was used, demonstrating that the OH reaction contributed to part of their formation. When OH radicals were scavenged, we speculated that PFA and PAA were mainly formed through ECIs isomerization and decomposition following the generation of HC(O)OO and $CH_3C(O)OO$ radicals. These radicals could further react with HO_2 forming peroxycarboxylic acids, which provided a plausible explanation for the higher yields of PFA and PAA observed in experiments with 2-butanol instead of

350   cyclohexane.

The yields of PFA and PAA had a positive relationship with water vapor concentration in all the cases, reaching the highest level at 90% RH. Model simulation results significantly underestimated the PAA formation and ignored the RH effect, indicating some missing pathways of PAA generation related to water. Because of the deficiency of PFA mechanism in MCM, we did not simulate the PFA formation. However, the similar variation

355   tendency of PFA and PAA yields could provide some evidence for the assumption that they might have similar formation mechanisms. In previous studies, the dominant formation pathway of PAA was considered to be the reaction of $CH_3C(O)OO$ radicals with HO_2 radicals (Groβ et al., 2014; Lightfoot et al., 1992; Winiberg et al., 2016), while the formation pathway of PFA had not been identified yet and was speculated to be the reaction of HC(O)OO radicals with HO_2 radicals (Liang et al., 2015). The rate constant of the RO_2 reaction with HO_2 was

360   investigated in a series of studies and was thought to be unaffected by water (Atkinson et al., 1999; Lightfoot et al., 1992; Tyndall et al., 2001; Wallington et al., 1992). Therefore, the increase of PFA and PAA yields with RH might be attributed to the promoting effect of water on HC(O)OO and $CH_3C(O)OO$ radical generation. Because

the highly reactive peroxides were easy to lose on the wall, chamber studies, which usually last for hours, might be hard to observe these peroxycarboxylic acids. In field observations, PAA is widely reported in remote and urban areas (Lee et al., 2000; Liang et al., 2013; Zhang et al., 2010), yet there are few reports on PFA existence in the atmosphere (Liang et al., 2015). The results here demonstrate that limonene ozonolysis could contribute to PFA and PAA formation, and although its high instability and reactivity makes PFA difficult to observe, it may have a short stay in the atmosphere.

3.5 Particulate peroxides and H_2O_2 generation

3.5.1 H_2O_2 evolution at low [O_3]/[limonene]

H_2O_2 generation from SOA in the aqueous phase is thought to be a possible way of producing H_2O_2 in cloud water or releasing H_2O_2 continuously after inhalation. Here, we quantitatively measured H_2O_2 generation from SOA produced in limonene ozonolysis in different cases. Considering that some unstable constituents would decay rapidly when SOA extract solution was kept at room temperature (298 K), we chose to keep the SOA solution at 4 ℃ to maintain the stability of the samples and prolong their storage time. In the low-ratio experiments, the concentration of total peroxides in solution nearly maintained stable in 48 h. The H_2O_2 concentration went through a brief period of increase and then became stable. After the aqueous H_2O_2 concentration reached a plateau, the amount of particulate H_2O_2 per particle mass formed in different cases, from 0% to 90% RH, was calculated and shown in Table 2. Regardless of whether the OH scavenger was used or not, SOA produced at higher RH was inclined to have a higher capacity of producing H_2O_2. When no OH scavenger was used, the lowest H_2O_2 content per particle mass was ~1.13 ng/μg at dry condition, and the highest value was ~2.45 ng/μg at 90% RH. When 2-butanol was used the changing trend of H_2O_2 generation resembled the condition without scavenger, and the minimum was ~1.33 ng/μg at 0% RH, while the maximum was ~2.89

ng/µg at 80% RH, which was a little higher than the value at 90% RH. It was interesting to note that, in the

385   presence of cyclohexane, the trend of H_2O_2 generation in SOA solution differed greatly from both of the above.

Even under dry condition, the H_2O_2 content per particle mass could reach ~3.22 ng/µg, and the maximum was

~4.63 ng/µg at 90% RH.

3.5.2 H_2O_2 evolution at high [O_3]/[limonene]

In the experiments of high [O_3]/[limonene], SOA produced with different scavengers was found to have

390   different rates of H_2O_2 generation in solution at 4 ℃, according to which we determined appropriate detection

frequency and total duration for the three kinds of SOA. SOA produced without OH scavenger had the lowest

rate of H_2O_2 generation, so an eight-day measurement result was reported here. For SOA produced in the

presence of 2-butanol, H_2O_2 generation was a little faster than the former and reached the limiting value within

six days. However, SOA produced with cyclohexane had a much faster rate than both of the above, and the H_2O_2

395   concentration in solution became stable within three days. For these three types of SOA, the total peroxides

concentration decreased slightly in the analysis duration, which was clarified in the Supplement.

It is obvious that SOA produced in the high-ratio experiments has greater ability of generating H_2O_2, and Fig. 4

shows the time profiles of H_2O_2 evolution of different kinds of SOA. When no OH scavenger was used, H_2O_2

concentration in solution rose constantly for six days then became stable. The limiting value of the H_2O_2

400   generation was influenced by RH, which was ~5.28 ng/µg at 0% RH, then increased gradually with increasing

RH until 80% RH (~ 14.45 ng/µg). In the experiments with 2-butanol, the H_2O_2 concentration kept rising in the

first four days, then it became stable. Under dry conditions, the limiting value of the H_2O_2 content was ~7.60

ng/µg, then it increased until 50% RH. SOA produced above 50% RH did not show any obvious difference, and

the highest H_2O_2 content was ~15.50 ng/µg. When cyclohexane was added, the H_2O_2 concentration in solution

405 increased rapidly in the first day then tended to become stable. The limiting value of H_2O_2 generation was also

affected by water vapor concentration, yet the promoting effect was not very significant, and no large difference

was observed above 30% RH. The limiting H_2O_2 content per particle mass was ~16.64 ng/μg at dry condition

and was ~30.00 ng/μg above 30% RH.

3.5.3 Different stabilities of particulate peroxides

410 According to the measurement results of the total peroxides and H_2O_2 in the SOA solution, the particulate

peroxides could be roughly divided into two categories: one is unstable that could decompose or hydrate to

generate H_2O_2, the other one is stable that could nearly remain unchanged in several days. Supposing that all the

peroxides contained one peroxy bond, the molar fractions of stable or unstable peroxides could be calculated as

follows:

$$\varphi_i = \frac{N_i}{N_{total}} \tag{1}$$

415 where φ_i is the molar fraction of stable or unstable peroxides in particles, N_i (mol) is the number of moles of

stable or unstable peroxides in particles, and N_{total} (mol) is the number of moles of total peroxides in particles.

At low $[O_3]/[limonene]$, the molar fraction of unstable peroxides in SOA was around 0.11–0.13 in the case of

adding no scavenger, which was similar to the case of adding 2-butanol. In the presence of cyclohexane, the

molar fraction of unstable peroxides would increase to 0.20–0.32. At high $[O_3]/[limonene]$, the molar fractions

420 of unstable peroxides in the case of adding no scavenger and adding 2-butanol were also similar, with both of

them in the range 0.13–0.25, yet this value would reach ~ 0.50 when cyclohexane was used. Detailed

information is provided in the Supplement. Model results showed that most peroxides produced in the reaction

were ROOH and R(O)OOH, in addition, some studies proposed that peroxyhemiacetals also made important

contribution to aerosol (Tobias and Ziemann, 2000; Tobias et al., 2000). The acetal reaction producing

425   peroxyhemiacetals is reversible, so part of the peroxyhemiacetals in SOA may have hydrolyzed to form some

peroxides. To investigate the stabilities of ROOH and R(O)OOH, we synthesized methyl hydroperoxide (MHP)

and ethyl hydroperoxide (EHP) to represent ROOH. PFA and PAA were used to represent R(O)OOH. All the

synthesis solutions were stored at 4 ℃, which was the same as the experimental condition. PFA and PAA were

found to decompose and generate H_2O_2 within several days, yet MHP and EHP maintained stable. Hence, we

430   speculate that peroxycarboxylic acids and peroxyhemiacetals might be the main components of unstable

peroxides in particles, and they contribute to H_2O_2 generation in the aqueous phase.

The amount of H_2O_2 generation at low $[O_3]/[limonene]$ measured here was comparable with the published value

of H_2O_2 produced from α-pinene SOA in solution (Li et al., 2016; Wang et al., 2011). However, at high

$[O_3]/[limonene]$, the H_2O_2 generation level increased significantly, which proved that the multi-generation

435   oxidation improved the formation of peroxycarboxylic acids and peroxyhemiacetals in particles. These results

demonstrated that SOA produced in limonene ozonolysis could behave as an important source of H_2O_2 in the

aqueous phase, and they also showed the difference between the SOA formed from single-unsaturated

monoterpene ozonolysis and double-unsaturated monoterpene ozonolysis.

3.5.4 The effect of OH scavenger

440   Prior researches provided evidence for OH radicals formation in ozonolysis experiments, and OH scavenger is

often used to avoid the disturbance of the OH reaction. Both 2-butanol and cyclohexane are commonly used OH

scavengers, and Chew and Atkinson (1996) showed that there was no difference in their abilities to scavenge

OH radicals. However, it should be noted that OH scavenger could convert OH radicals into a mixture of

hydroperoxy (HO_2) and alkylperoxy (RO_2) radicals and a higher $[HO_2]/[RO_2]$ is observed when 2-butanol is

445   used (Docherty and Ziemann, 2003; Jonsson et al., 2008a). As shown in the Sect. 3.6.1, the fact that both at low

and high $[O_3]/[limonene]$, the SOA yield was higher with 2-butanol than cyclohexane suggested that the

increase in HO_2 concentration promoted the SOA formation. This result is consistent with that suggested by Keywood et al. (2004) who observed the higher $[HO_2]/[RO_2]$ resulting in higher SOA yield in cyclohexene ozonolysis, while Docherty and Ziemann (2003) showed that increased $[HO_2]/[RO_2]$ inhibited aerosol formation in β-pinene ozonolysis.

Some studies demonstrated that acid and peroxide products were sensitive to OH scavengers (Ahmad et al., 2017; Henry and Donahue, 2011; Ma et al., 2008). Here through investigating the stabilities of particulate peroxides formed in different conditions, we showed that the choice of OH scavenger would influence the types of particulate peroxides produced in the reaction. In the Sect. 3.5.3, we suggest that peroxycarboxylic acids and peroxyhemiacetals be the main components of unstable peroxides in particles. The peroxyhemiacetals formed by heterogeneous reactions of peroxides and aldehydes could dissociate into these species, yet the formation of peroxyhemiacetals was not considered to be affected by the OH scavenger since both the peroxides and carbonyl compounds observed in the gas phase did not show large differences when the OH scavenger changed. We speculate that the impact of OH scavenger on particulate peroxides could be mainly attributed to the formation of peroxycarboxylic acids under different $[HO_2]/[RO_2]$. Keywood et al. (2004) indicated that the OH scavenger impacted the HO_2 reaction with acylperoxy radicals forming acid and peracid products, which were considered to have low volatility. With higher HO_2 concentration, the SOA yield in the presence of 2-butanol was higher than that in the presence of cyclohexane, since the HO_2-acylperoxy reactions contributed to some low-volatility products. However, the RO_2 radicals formed from cyclohexane may also participate in reactions and help produce more unstable peroxycarboxylic acids which could partition into the particle phase.

3.6 Contribution of peroxides to SOA

3.6.1 SOA formation

The yield of SOA produced from limonene ozonolysis in different cases ranging from 0% to 90% RH was measured. The SOA yield is defined as the ratio of aerosol mass concentration to the mass concentration of limonene consumed, as given by Eq. (S4) (in the Supplement). In the six different cases, SOA yield was determined to be unaffected by water, yet it showed strong dependence on the reactants ratio and the use of OH scavenger, as shown in Table 3. Whether at low or high $[O_3]$/[limonene], the case of no OH scavenger had the highest SOA yield, which suggested the effect of OH reaction on aerosol formation, and the order of SOA yield was no scavenger > 2-butanol > cyclohexane. The case with the highest SOA yield was high $[O_3]$/[limonene] without OH scavenger (0.511 ± 0.097), while the lowest SOA yield was detected at low $[O_3]$/[limonene] with cyclohexane (0.288 ± 0.038).

3.6.2 Peroxides mass fraction

In this study, we used the iodometric method to analyze the particle-phase total peroxides content in limonene ozonolysis. The iodometric method is commonly regarded as a standard method for detecting total peroxides and it could almost quantify all kinds of peroxides (Bonn et al., 2004; Jenkin, 2004). The mass fraction used here is defined as the ratio of particulate peroxides mass to SOA mass, as given in Eq. (S5) (in the Supplement). The average molecular weight of peroxides in particles is assumed to be 300 g/mol, which is estimated to be slightly less than the molecular weight of peroxyhemiacetals, and this value has been used to calculate the mass of particulate peroxides in some studies (Docherty et al., 2005; Nguyen et al., 2010; Surratt et al., 2006). The mass fractions of peroxides in the six cases from 0% to 90% RH are summarized in Table 3. In each case, the peroxides mass fraction in SOA increased slightly with RH, and there was a significant difference between the

low-ratio and high-ratio experiments, while the there was little effect of the OH scavenger. The mass fraction of peroxides at low $[O_3]/[limonene]$ was usually below 0.2, however, at high $[O_3]/[limonene]$ the peroxides mass fraction increased to 0.4–0.6. Docherty et al. (2005) reported that the peroxides mass fraction in α-pinene SOA was ~0.47, while for β-pinene SOA the fraction was ~0.85. Li et al. (2016) observed that peroxides could account for ~0.21 in α-pinene SOA. We are the first to report the mass fraction of peroxides in SOA derived from limonene ozonolysis, highlighting the important role of organic peroxides in SOA composition, especially when multi-generation oxidation happens.

3.6.3 Peroxides partitioning

The molar yields of HMW peroxides in both gas phase and particle phase were determined in different cases. There was little effect of RH and OH scavenger, while at high $[O_3]/[limonene]$ the HMW peroxides yield increased by ~ 10% in contrast with the yield at low $[O_3]/[limonene]$. Furthermore, the gas-phase HMW peroxides and particle-phase HMW peroxides were discussed separately, which was shown in Fig. 5. For gas-phase HMW peroxides, the molar yield showed a decreasing tendency with increasing RH, which was obvious at high $[O_3]/[limonene]$. For particle-phase HMW peroxides, the molar yield showed a slight increasing dependence on RH and their yield was promoted significantly by the degree of oxidation. A possible explanation for the RH effect was that the water content in aerosol would increase when RH increased, and this promoted the uptake of some compounds into particles. However, some researches proved that the effect of water on the partitioning of organic products seemed to be small especially when no inorganic seed particles were used (Jonsson et al., 2006, 2008a, b). Thus, we propose that the promoting effect of water on particulate HMW peroxides can be mainly attributed to the effect of water participating in some gas-phase and particle-phase reactions resulting in the formation of more low-volatility peroxides. The contribution of multi-generation

oxidation products to SOA formation in limonene ozonolysis had been stated by some studies (Hoffmann et al., 1997; Ng et al., 2006), and the results here indicate that organic peroxides might account for a considerable

510   proportion of those products. Multi-generation oxidation helps produce more low-volatility peroxides in the gas phase that could partition into particles, What's more, it accelerates the occurrence of some heterogeneous reactions by providing more reactive species.

3.7 Contribution of carbonyls to SOA

3.7.1 Carbonyls formation

515   Carbonyl compounds including HACE, FA, AA, ACE, GL, and MGL were detected in the reaction. Unlike peroxides, whose generation often showed obvious increasing dependence on RH, only FA yield would increase with RH and the production of other carbonyls were not significantly affected by water. The yields of HACE, AA, ACE, GL, and MGL in different cases were summarized in Table 4. The fact that HACE formed without OH scavenger but did not form when cyclohexane was used indicated that HACE might be generated from OH

520   reactions. As for AA and ACE, both of their yields in the presence of cyclohexane were found to be lower than the case without scavenger, especially at low $[O_3]/[limonene]$. This suggested that OH reaction contributed to a potion of their formation and the endocyclic DB ozonolysis did not tend to generate AA and ACE. It was speculated that these three kinds of carbonyls could also generate from the reaction of 2-butanol or HO_2 radicals promoted their formation, so the presence of 2-butanol increased their yields. As regards GL and MGL, RH and

525   OH scavenger did not have a large influence, and both of their yields at high $[O_3]/[limonene]$ were higher than the condition of low $[O_3]/[limonene]$. The generation of FA had a positive dependence on RH, which was illustrated in the Supplement and the total variation range could be found in Table 4.

3.7.2 Experimental and theoretical partitioning coefficients

A parameter that has been used widely to describe the partitioning feature of a compound is described as follows (Pankow, 1994; Pankow and Bidleman, 1992):

$$K_{p,i} = \frac{F_i/TSP}{A_i} \qquad (2)$$

where $K_{p,i}$ (m^3 μg^{-1}) is the partitioning coefficient of compound i, TSP (μg m^{-3}) is the concentration of total suspended particulate matter, F_i (μg m^{-3}) and A_i (μg m^{-3}) are the particulate and gaseous concentrations of compound i, respectively. The measured partitioning coefficients of FA, AA and ACE were on the magnitude of 10^{-5} m^3 μg^{-1}, and the partitioning coefficients of HACE, GL, and MGL were on the magnitude of 10^{-4} m^3 μg^{-1}.

Theoretical gas-particle partitioning coefficients of these compounds were calculated using the absorption equilibrium equation defined by Pankow (1994), which has been used widely to estimate the ability of a substance to partition into the particulate phase and to predict SOA yield in models (Griffin et al., 1999; Hohaus et al., 2015; Odum et al., 1996; Yu et al., 1999):

$$K_{p,i} = \frac{760\, R\, T\, f_{om}}{MW_{om}\, 10^6\, \zeta_i\, p_{L,i}^0} \qquad (3)$$

where R is the ideal gas constant (8.206×10^{-5} m^3 atm mol^{-1} K^{-1}), T is the temperature (K), f_{om} is the mass fraction of TSP that is the absorbing organic material (om) and is equal to 1 here, MW_{om} (g mol^{-1}) is the mean molecular weight of the om phase, which is estimated to be 130 g mol^{-1} in this study, ζ_i is the activity coefficient of compound i in the om phase, which is assumed to be unity, and $p_{L,i}^0$ (Torr) is the vapor pressure of compound i, which is predicted by the method of Moller et al. (2008). The calculated gas-particle partitioning coefficients of HACE, FA, AA, ACE, GL, and MGL were 3.972×10^{-8} m^3 μg^{-1}, 2.096×10^{-11} m^3 μg^{-1}, 1.624×10^{-10} m^3 μg^{-1}, 6.192×10^{-10} m^3 μg^{-1}, 5.476×10^{-10} m^3 μg^{-1}, and 1.246×10^{-9} m^3 μg^{-1}, respectively, at 298K. The experimental K_p value of HACE was about 10,000 times larger than the theoretical value, and for other

carbonyls, the experimental K_p was about 100,000 times larger than the theoretical value. The gap between the experimental K_p and predicted K_p of GL and MGL we estimated was comparable with the results of Healy et al. (2008, 2009), but it was higher than that of Ortiz et al. (2013). The fact that the gas-particle partitioning coefficients of carbonyls observed were much higher than the theoretical values indicated that carbonyl compounds made a more important contribution to SOA formation than estimated. Although the partitioning coefficients measured in experiments showed huge difference with the calculated values, a relationship between the measured K_p and the vapor pressure of carbonyl compounds was observed. Figure 6 shows the dependence of the measured $lg(K_p)$ and predicted $lg(K_p)$ on $lg(p^0)$, and their linear fitting curves. The slope of the linear fitting equation of the predicted $lg(K_p)$ versus $lg(p^0)$ was -0.964, and R^2 was 0.998. The slope of the linear fitting equation of measured $lg(K_p)$ versus $lg(p^0)$ was -0.484, and R^2 was 0.750, which indicated that the $lg(K_p)$ of carbonyls observed in laboratory also had a negative correlation with $lg(p^0)$. A plausible explanation for the large difference between the measured and predicted K_p was that carbonyl compounds were easy to polymerize and react with other species on particles, resulting in these carbonyls existing in the form of hydrates and oligomers in SOA (Corrigan et al., 2008; Hastings et al., 2005; Kroll et al., 2005; Volkamer et al., 2007). The hydrates and oligomers of carbonyls have much lower vapor pressures than their precursors, and they could reversibly return to their carbonyl monomers during analysis (Healy et al., 2008; Ortiz et al., 2013; Toda et al., 2014).

4 Conclusions and implications

This work reports on an experimental study of the oxidation regime and SOA composition in limonene ozonolysis with respect to the roles of different DBs, radicals, and water. A series of oxidizing products including SCIs, OH radicals, and peroxides in both gaseous and particulate phases were detected. Based on the

variation of H_2O_2 and HMHP generation on RH, the importance of the reaction of limonene SCIs with water was confirmed and the yield of SCIs could be estimated as ~0.24 for endocyclic DB and ~0.43 for exocyclic DB.

570   OH yields of endocyclic and exocyclic DBs were indirectly determined to be ~0.65 and ~0.24, demonstrating the different reaction mechanisms of different DBs in limonene ozonolysis. The formation of two peroxycarboxylic acids PFA and PAA was observed, and their high yields were first reported in alkene ozonolysis. The yields of PFA and PAA increased with RH and the degree of oxidation, showing the effect of water and the exocyclic DB oxidation on their formation. The H_2O_2 generation from SOA in solution provided

575   evidence for the ability of SOA to contribute oxidants in the aqueous phase. And through investigating the stabilities of particulate peroxides formed under different conditions, we showed that the choice of the OH scavenger would influence the types of particulate peroxides produced in the reaction. The partitioning behaviors of peroxides and carbonyls were discussed in this paper, and the results showed their importance to SOA formation. Particulate peroxides could account for less than 0.2 in limonene SOA at low $[O_3]/[limonene]$

580   and 0.4–0.6 at high $[O_3]/[limonene]$, which proved the important role of peroxides in SOA composition especially when multi-generation oxidation happened. Due to their ability to polymerize and react with other species on particles, the carbonyl compounds often exist in SOA in forms of hydrates and oligomers whose vapor pressures are much lower than that of their precursors. The partitioning coefficients of carbonyls observed in laboratory were always several orders of magnitude higher than the theoretical values, indicating that their

585   contribution to SOA was higher than estimated. Limonene is unique in many aspects because of its different DBs, suggesting that terpenes with multiple double bonds may have more complex effects on the atmosphere than previously thought. The results demonstrated the effect of the degree of oxidation, radical chemistry, and water on limonene SOA composition, while some particulate products especially peroxides should be studied further. The implications of limonene chemistry in the atmosphere are close with the multi-generation oxidation

590   that occurs on its different DBs, and whether the phenomena could be observed in other terpenes oxidation that

contain more than one DB needs more laboratory evidence.

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Table 1. Experimental Conditions.

Exp.	[Limonene] (ppbv)	[O$_3$] (ppbv)	OH Scavenger	[OH Scavenger] (ppmv)	RH (%)
L(No-sca)	280	500	—	—	0–90
L(2-But)	280	500	2-butanol	350	0–90
L(C-hex)	280	500	cyclohexane	420	0–90
H(No-sca)	183	19000	—	—	0–90
H(2-But)	183	19000	2-butanol	350	0–90
H(C-hex)	183	19000	cyclohexane	420	0–90

Note: L, low [O$_3$]/[limonene]; H, high [O$_3$]/[limonene]; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane; RH,

990   relative humidity.

Table 2. H_2O_2 generation per particle mass (ng/ug) in SOA formed with different OH scavengers in the relative humidity (RH) range of 0%−90% at low $[O_3]$/[limonene] ratio.

	0% RH	10% RH	30% RH	50% RH	70% RH	80% RH	90% RH
No-sca	1.13±0.22	1.42±0.40	1.53±0.28	1.88±0.16	2.24±0.42	2.26±0.44	2.45±0.48
2-But	1.33±0.15	1.56±0.18	1.77±0.12	2.02±0.65	2.55±0.43	2.89±0.42	2.66±0.57
C-hex	3.22±0.52	3.95±0.43	4.12±0.40	4.22±0.33	4.63±0.24	4.26±0.33	4.63±0.96

Note: No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

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Table 3. The SOA yield and mass fraction of particulate peroxides at low or high $[O_3]/[limonene]$ ratio in the presence or absence of OH scavenger from 0% to 90% relative humidity (RH).

	L(No-sca)	L(2-But)	L(C-hex)	H(No-sca)	H(2-But)	H(C-hex)
0% RH	0.065±0.006	0.101±0.009	0.087±0.011	0.401±0.016	0.502±0.008	0.477±0.010
10% RH	0.091±0.010	0.124±0.013	0.113±0.009	0.436±0.009	0.534±0.009	0.502±0.013
30% RH	0.125±0.010	0.147±0.011	0.143±0.015	0.458±0.020	0.553±0.015	0.506±0.011
50% RH	0.149±0.007	0.174±0.011	0.161±0.016	0.466±0.016	0.571±0.009	0.512±0.007
70% RH	0.155±0.009	0.178±0.009	0.169±0.014	0.486±0.023	0.576±0.010	0.503±0.011
80% RH	0.169±0.013	0.189±0.010	0.189±0.012	0.492±0.015	0.580±0.013	0.502±0.014
90% RH	0.156±0.010	0.183±0.013	0.189±0.013	0.492±0.017	0.580±0.018	0.506±0.016
SOA Yield	0.379±0.039	0.337±0.048	0.288±0.038	0.511±0.097	0.479±0.044	0.401±0.068

Note: L, low $[O_3]/[limonene]$; H, high $[O_3]/[limonene]$; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

1000 **Table 4.** Yields (%) of carbonyls at low or high $[O_3]/[\text{limonene}]$ ratio in the presence or absence of OH scavenger.

	HACE	FA	AA	ACE	GL	MGL
L(No-sca)	2.04±0.48	7.02±0.90–10.58±0.94	1.32±0.24	0.22±0.15	0.89±0.25	0.56±0.34
L(2-But)	3.94±0.55	4.90±0.86–7.77±0.86	3.98±0.60	0.35±0.18	0.69±0.25	0.55±0.44
L(C-hex)	—	5.21±0.66–8.25±0.55	—	—	0.81±0.24	0.67±0.56
H(No-sca)	4.45±0.52	13.11±0.63–27.00±1.56	2.16±0.84	0.79±0.22	1.33±0.41	1.35±0.61
H(2-But)	10.15±2.11	11.03±0.77–23.33±0.62	7.86±1.32	1.00±0.28	1.25±0.36	1.31±0.21
H(C-hex)	—	10.80±1.28–23.32±1.21	1.89±1.22	0.60±0.47	1.25±0.51	1.23±0.22

Note: —, below detection limit; L, low $[O_3]/[\text{limonene}]$; H, high $[O_3]/[\text{limonene}]$; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane; HACE, hydroxyacetone; FA, formaldehyde; AA, acetaldehyde; ACE, acetone; GL, glyoxal; MGL, methylglyoxal.

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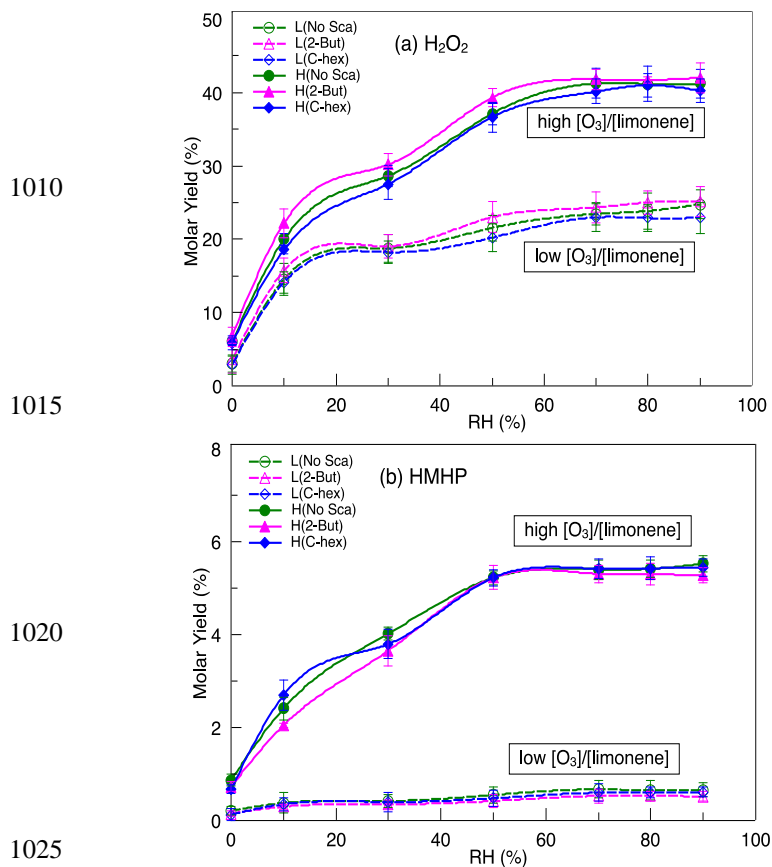
Figure 1. Dependence of (a) H_2O_2 yield and (b) HMHP yield on relative humidity (RH) at low or high $[O_3]/[limonene]$ ratio

in the presence or absence of OH scavenger (2-butanol or cyclohexane). H_2O_2, hydrogen peroxide; HMHP, hydroxymethyl

hydroperoxide; L, low $[O_3]/[limonene]$; H, high $[O_3]/[limonene]$; No-sca, none scavenger; 2-But, 2-butanol; C-hex,

    cyclohexane.
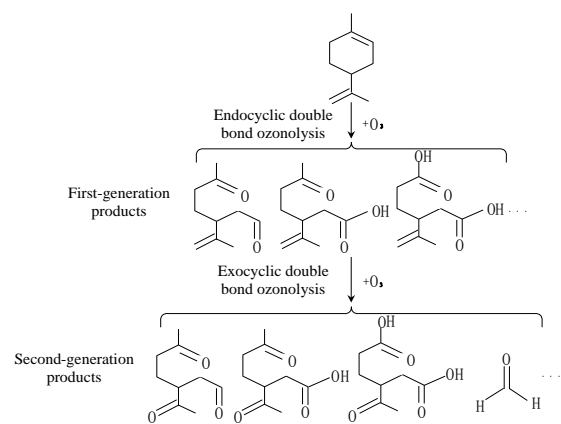
Figure 2. The reaction scheme of endocyclic and exocyclic double bonds in limonene ozonolysis

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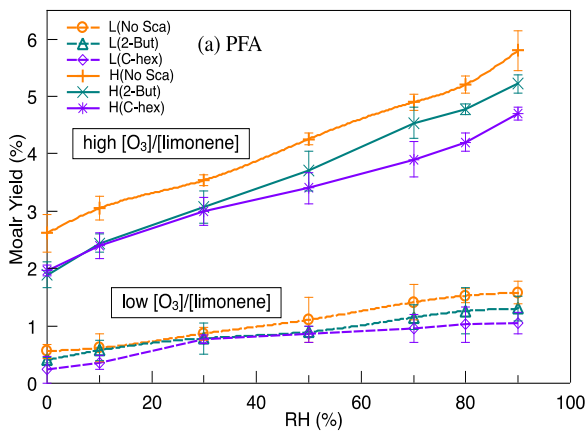
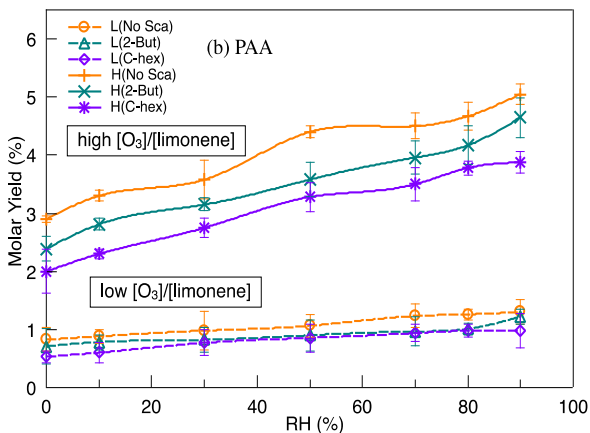
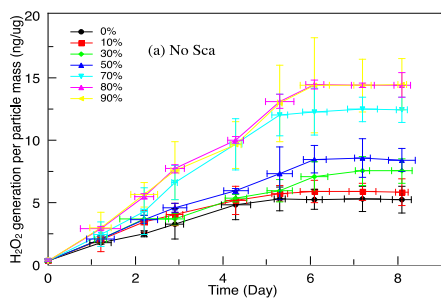
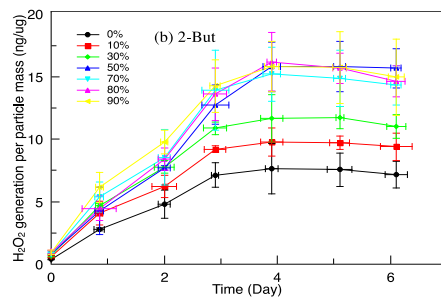
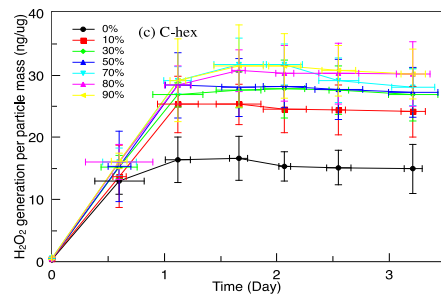
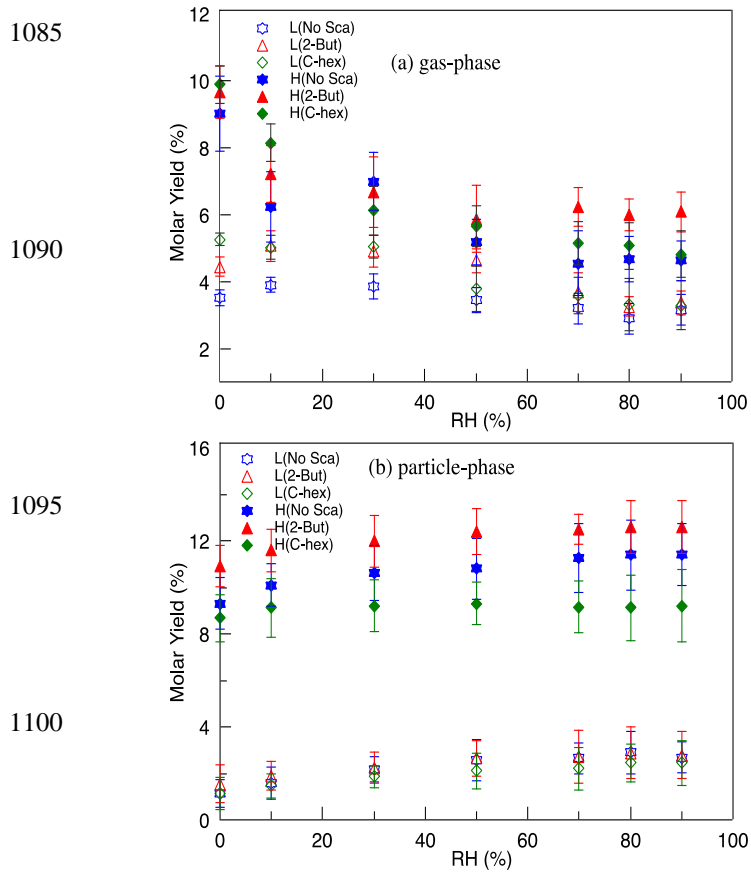


Figure 3. Dependence of (a) PFA yield and (b) PAA yield on relative humidity (RH) at low or high [O_3]/[limonene] ratio in the presence or absence of OH scavenger (2-butanol or cyclohexane). PFA, peroxyformic acid; PAA, peroxyacetic acid; L, low [O_3]/[limonene]; H, high [O_3]/[limonene]; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

1080 **Figure 4.** Time profiles of H$_2$O$_2$ evolution per particle mass of different SOA formed (a) without OH scavenger, (b) with 2-butanol, and (c) with cyclohexane in the relative humidity (RH) range of 0–90% at high [O$_3$]/[limonene] ratio. H$_2$O$_2$, hydrogen peroxide; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

1105 **Figure 5.** The variation of (a) gas-phase and (b) particle-phase high-molecular-weight peroxides molar yields with relative

humidity (RH) at low or high $[O_3]/[limonene]$ ratio in the presence or absence of OH scavenger (2-butanol or cyclohexane).

L, low $[O_3]/[limonene]$; H, high $[O_3]/[limonene]$; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

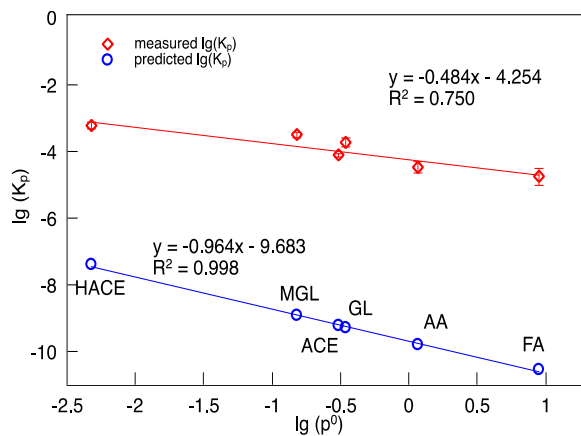



Figure 6. The relationship of measured and predicted partitioning coefficients ($m^3\ \mu g^{-1}$) versus vapor pressure (atm) of carbonyls produced in limonene ozonolysis. K_p, partitioning coefficients, p^0, vapor pressure, HACE, hydroxyacetone; FA, formaldehyde; AA, acetaldehyde; ACE, acetone; GL, glyoxal; MGL, methylglyoxal.