August 2, 2018

ACP Editor

Dear Prof. Markus Ammann,

Enclosed please find our revised manuscript entitled "***The oxidation regime and SOA composition in limonene ozonolysis: Roles of different double bonds, radicals, and water***", revised supplement and two responses to the anonymous referees #1 and #2. We gratefully thank for the two reviewers' constructive suggestions to help us improve the manuscript.

The major revisions are specified as follows:

1. The title has been changed into "The oxidation regime and SOA composition in limonene ozonolysis: Roles of different double bonds, radicals, and water". (Previous title is "Understanding the oxidants transition and SOA property in limonene ozonolysis: Role of different double bonds, radical chemistry, and water".)
2. All of the introductory materials shown in the results have been moved to the introduction, and the introduction has been restructured and shortened.
3. The discussion of SCIs generation is improved to be more complete.
4. An explanation for the impact of different OH scavengers is added in the text.
5. The detection methods and the experimental conditions are clarified more clearly.
6. A scheme of endocyclic and exocyclic double bonds of limonene reaction with ozone is added in the text.
7. A series of calculation equations and diagrams of the reactor and collector are added in the Supplement.
8. The effect of high reactants concentrations is discussed in the Supplement.

Detailed changes made in the manuscript can be seen in the marked-up version in this response.

Thanks for your time.

Sincerely yours,

Zhongming Chen and co-authors

Response to Reviewer 1

We gratefully thank you for your constructive comments and thorough review. Below are our responses to your comments.

(Q=Question, and A=Answer)

General Comments:

Q1. The introduction is far too long. Also, significant quantities of introductory material appear in the results (lines 264-285 and 325-340). This material should all appear in the intro, which itself needs to be written more concisely. The authors should also try and use better paragraph structure in places and use shorter paragraphs in general.

A: Thanks for your suggestion. All of the introductory materials shown in the Sect. 3.2, Sect. 3.3, and Sect. 3.4 are now moved to the introduction. The introduction has been restructured and shortened.

Q2. Another general point is that the methods do not provide enough detail. For each technique discussed, the authors should provide detection limits, precision and accuracy.

A: We follow the advice of the reviewer. The detection limits, precision, and accuracy have been provided for the methods of detecting peroxides and carbonyls in the revised text. We also provide the diagrams of the flow tube reactor, the coil collector, and the Horibe tube in the Supplement

Q3. Throughout, there is a tendency to talk about things being higher, lower or different. It would be better if such expressions were quantified where possible.

A: Thanks for your suggestion. We have checked our manuscript and changed such expressions.

Q4. Finally, the experiments are carried out under conditions far removed from most ambient atmospheres, both in terms of concentrations of limonene and ozone, but also temperature (277K). Given this, I was missing a statement of the applicability of these results to real atmospheric conditions, or indeed as alluded to in the manuscript, indoor environments.

A: We regret that we did not clarify well enough. All of our experiments were conducted at 298 K as stated in the Sect. 2.2. The temperature of 277 K was just used to keep the SOA solution in order to maintain the stability of samples and prolong their storage time. We have explained this in the revised manuscript. As for the concentrations of reactants used in our experiments, the concerns of the reviewer are reasonable. To get enough products for analysis in a short reaction time, both of the concentrations of limonene and ozone in this study were obviously higher than those in the real atmospheric conditions, which might have influence on the gas-phase and particle-phase chemistry. So the effects of the reactants concentrations on the experimental results are discussed below and this part is now added in the Supplement.

A major impact of the high concentrations of reactants is the increased RO_2 concentration. In recent years many studies reported that the autoxidation processes formed highly oxidized RO_2 radicals, which reacted with HO_2 and other RO_2 radicals forming highly oxidized multifunctional organic compounds (HOMs) (Jokinen et al., 2014; Richters et al., 2016a, b). The production of HOMs is controlled by two competing processes, i.e., RO_2 autoxidation vs. RO_2 reaction with HO_2 and other RO_2 radicals. Zhang et al. (2015) found that at low α-pinene levels, the longer lifetime of RO_2 radicals favored the isomerization pathways and consequently led to enhanced ELVOC dimers production. They estimated that the corresponding lifetime of RO_2 radicals decreased by less than an order of magnitude when the initial α-pinene mixing ratio increased from 10 ppbv to 150 ppbv, which was not sufficient to perturb the dynamics of overall RO_2 chemistry. In our experiments, where the limonene concentration was below 200 ppbv, we speculated that although the RO_2 chemistry was affected to some extent it would not bring huge influence on the results. When the RO_2 concentration is high, the reactions of SCIs and RO_2 radicals might happen in the system (Sadezky et al., 2008). Zhao et al. (2015) found that the reactions of SCIs and RO_2 radicals played a key role in particle formation in *trans*-3-hexene ozonolysis, while for large alkenes such as terpenes and sesquiterpenes such reactions might be unimportant. Thus, although the concentrations of reactants used in this study were higher than that in the real atmospheric condition, the SCIs + RO_2 reactions did not have a huge effect on the reaction system. The SOA yield of limonene ozonolysis observed in this study was in the range of the values reported before (Ahmad et al., 2017; Chen and Hopke, 2010; Pathak et al., 2012). It is true that when the mixing ratios of reactants are high, the gas-particle partitioning processes of semi-volatility and low-volatility products are promoted resulting in higher SOA yield, yet we think that it may not have great impact on the representativeness of the products we investigated in particles.

Specific Comments:

Q6. I am not going to list every grammatical error, but I will mention a few here. I think that when the authors talk about ’SOA property’ they mean SOA composition. This phrase should be replaced throughout with something more informative. Similarly, I am not clear what they mean by ’oxidants transition’. Again, clarification is needed.

A: Thanks for your suggestion and we have changed “SOA property” to “SOA composition” throughout the text. Besides, “oxidants transition” is replaced by “oxidation regime” in the revised manuscript. In this study, the objective of investigating the oxidation regime of alkene ozonolysis is to understand the formation of the oxidizing products in limonene ozonolysis. These compounds, including OH radicals, stabilized Criegee intermediates, and peroxides, are critical to atmospheric oxidation processes since they own the power of oxidizing other species. We clarify this in the revised introduction.

Q7. ‘Particulate unstable peroxides’ is an awkward expression.

A: We have changed that to “unstable peroxides in particles”.

Q8. Line 64 – ‘eaters’ - esters?

A: Yes, we have revised it.

Q9. Line 542: Assume these are Kp coefficients - should specify. I don't understand the explanation for the finding that Kp is much bigger than estimated, line 556 (and lines 579-580). The language needs to be improved. What impact could experimental conditions have (low T and high precursor concentrations).

A: We have specified that these are the calculated gas-particle partitioning coefficients. As for the explanation for the finding that the measured K_p is much bigger than the predicted K_p, we have clarified it better in the revised text. A plausible explanation for the large difference between the measured K_p and the predicted K_p was that carbonyl compounds were easy to polymerize and react with other species on particles, resulting in that these carbonyls existed in forms of hydrates and oligomers (Corrigan et al., 2008; Hastings et al., 2005; Kroll et al., 2005; Volkamer et al., 2007). The hydrates and oligomers of carbonyls have much lower vapor pressures than their precursors, and they could reversibly return to their carbonyl monomers during analysis (Healy et al., 2008; Ortiz et al., 2013; Toda et al., 2014). The temperature used here was 298 K and the effect of precursor concentrations was analyzed. When the mixing ratios of reactants are high, the gas-particle partitioning processes of semi-volatility and low-volatility products are promoted resulting in higher SOA yield. However, when we calculated the gas-particle partitioning coefficients, the effect of the concentration of total suspended particulate matter was taken into consideration as shown by Eq. (2), so the higher SOA concentration caused by higher precursor concentrations would not impact the partitioning behaviors of carbonyls. Another effect of the high precursor concentrations was that it might promote the reactions of carbonyl compounds on particles, yet the concentrations of most carbonyls detected in our reactions were usually several ppbv, which were just a little higher than those in the real atmosphere, so we speculated that this impact on the partitioning coefficients was also limited.

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Response to Reviewer 2

We gratefully thank you for your constructive comments and thorough review. Below are our responses to your comments.

(Q=Question, and A=Answer)

General Comments:

Q1. It seems that “SOA property” should be replaced by SOA composition.

A: We have revised it.

Q2. It seems that “oxidation transition” should be replaced by “oxidation regime” or something similar. Do you mean to indicate the effect of adding OH scavenger? Please clarify.

A: Thanks for your suggestion and “oxidation regime” is used in the revised manuscript. In this study, the objective of investigating the oxidation regime of alkene ozonolysis is to understand the formation of the oxidizing products in limonene ozonolysis. These compounds, including OH radicals, stabilized Criegee intermediates, and peroxides, are critical to atmospheric oxidation processes since they own the power of oxidizing other species. We clarify this in the revised introduction.

Q3. There are a number of places where introductory material shows up in the results and discussion, such as the historical perspective on OH formation in section 3.3. Please move this material to the introduction. The introduction itself needs a much clearer/more logical flow of ideas.

A: We follow the advice of the reviewer and all of the introductory materials shown in the Sect. 3.2, Sect. 3.3, and Sect. 3.4 are now moved to the introduction. The introduction has been restructured and shortened.

Q4. Add a clearer discussion/explanation of why the two OH scavengers had different effects.

A: The reviewer concerns the effect of two OH scavengers hence the following content has been added into the Sect. 3.5.4 of the revised text.

Researches before have provided evidence for OH radicals formation in ozonolysis experiments and OH scavenger is often used to avoid the disturbance of OH reaction. 2-butanol and cyclohexane used here are both commonly used OH scavengers, and Chew and Atkinson (1996) showed that there was no difference in their abilities to scavenge OH radicals. However, it should be noted that OH scavenger could convert OH radicals into a mixture of hydroperoxy (HO_2) and alkylperoxy (RO_2) radicals and higher $[HO_2]/[RO_2]$ is observed when 2-butanol is used (Docherty and Ziemann, 2003; Jonsson et al., 2008). As shown in the Sect. 3.6.1, the fact that both at low and high $[O_3]$/[limonene], the SOA yield was higher with 2-butanol than with cyclohexane suggested that the increase of HO_2 concentration promoted the SOA formation. This result is consistent with that suggested by Keywood et al. (2004) who observed the higher $[HO_2]/[RO_2]$ resulting in higher SOA yield in cyclohexene ozonolysis, while Docherty and Ziemann (2003) showed that increased $[HO_2]/[RO_2]$ inhibited aerosol formation in β-pinene ozonolysis.

Some studies demonstrated that acid and peroxide products were sensitive to OH scavengers (Ahmad et al., 2017; Henry and Donahue, 2011; Ma et al., 2008). Here through investigating the stabilities of particulate peroxides formed in different conditions, we showed that the choice of OH scavenger would influence the types of particulate peroxides produced in the reactions. In the Sect. 3.5.3, we suggest that peroxycarboxylic acids and peroxyhemiacetals be the main components of unstable peroxides in particles. The peroxyhemiacetals formed by heterogeneous reactions of peroxides and aldehydes could dissociate into these species, yet the formation of peroxyhemiacetals was not supposed to be affected by OH scavenger since both of the peroxides and carbonyl compounds observed in the gas phase did not show large differences when OH scavenger changed. We speculate that the impact of OH scavenger on particulate peroxides could be mainly attributed to the formation of peroxycarboxylic acids under different $[HO_2]/[RO_2]$. Keywood et al. (2004) indicated that the OH scavenger impacted the HO_2 reaction with acylperoxy radicals forming acid and peracid products, which were considered to have low volatility. With higher HO_2 concentration, the SOA yield in the presence of 2-butanol was higher than that in the presence of cyclohexane, since the HO_2-acylperoxy reactions contributed to some low-volatility products. However, the RO_2 radicals formed from cyclohexane may also participate into reactions and help produce more unstable peroxycarboxylic acids which could partition into the particle phase.

Q5. Add a clearer discussion of measurement techniques, particular the coil collector, with diagrams of the flow reactor and coil collector in the supplement.

A: We have better explained the measurement techniques in the revised manuscript. The coil collector is around 30 cm long and its effective length is about 100 cm. The coil is similar with that used in earlier studies (Grossmann et al., 2003; Sauer et al., 1996, 1997) and is controlled at a temperature of 4 °C with H_3PO_4 solution (pH 3.5) serving as the rinsing solution. The diagrams of the flow tube reactor, the coil collector, and the Horibe tube are added in the Supplement.

Q6. The differences in yields are often overstated or exaggerated. For example, the peroxide mass fraction was determined in the three scavenger cases and discussed as if there were significant differences among these cases (Lines 484-487). But in fact, the differences were fairly small. Please check over all comparisons in the manuscript and only state differences if they can be supported statistically, otherwise characterize them as similar.

A: We have revised it and other comparisons in the text are also checked.

Q7. A theme of the paper is the different effects from each of the two double bonds in limonene reacting with ozone. But your discussion of the ozonolysis of the endo and exo double bonds (e.g, Lines 320-323) is misleading, because the ozonolysis of the endo-DB will generally precede ozonolysis of the exo-DB. Your experiments can only isolate the ozonolysis of the endo-DB due to the discrepancy in reaction rates, which you do state. The ozonolysis of what was originally the exo-DB may not occur on the exo-DB of limonene, but rather on the remaining DB in an ozonolysis product of limonene. This is clearly stated in the final paragraph of Herrmann et al. “It should be noted that the measured OH-radical yield of the second double bond is the sum of all possible products formed from the reaction of ozone with the first double bond of the monoterpene.”(1)

A: We regret that in the previous version we did not clarify well enough. We have added an explanation in Sect. 3.2 for that. The SCIs yield of exocyclic DB ozonolysis estimated here was not for a specific product formed from endocyclic DB ozonolysis, but the sum of first-generation products with a remaining double bond.

Q8. Too often the past tense is used incorrectly. For example, line 288 should read, “Figure 3 shows . . .”

A: We have revised it and other similar mistakes are also corrected.

Q9. Calculated values are rarely supported by equations that clearly show how the calculation was done (however simple they may be). Add equations, either in the main text or supplement, that show how all yields are calculated.

A: We follow the advice of the reviewer and in the Supplement a summary of the calculation equations have been added, which includes the equations to calculate the wall loss fraction, the molar yield of products, the molar yield of SCIs, the SOA yield, and the mass fraction of particulate peroxides.

Specific Comments:

Q10. Figures 1 and 2. Neither of these needs to be in the main manuscript, both should be moved to supplemental. You should replace these with a scheme that indicates what chemistry you are investigating, particularly to show that the second DB oxidized is part of the products from the initial ozonolysis (such as Fig.2 from Herrmann et al.). (1)

A: Figure 1 and Figure 2 have been moved to the Supplement and a scheme of endocyclic and exocyclic double bonds of limonene reaction with ozone (Figure 2) is added in the revised manuscript.

Q11. Figure 3. The different markers are hard to distinguish. Please change the markers you use to make it easier to determine what experiment each line corresponds to. You might also label the two groups of lines (low and high ratios of ozone to limonene).

A: Thanks for your suggestion. We have changed the markers and labeled the two groups of lines in Figure 1 and Figure 3 in the revised text.

Q12. Figure 6. The lines are not helpful, please use just markers for each data point.

A: We have revised it.

Q13. Lines 75-79 The ability of SCI to react with water is grossly overstated, even though it may be an important source of peroxides in the case of limonene. Review the discussions and add references to publications such as Long et al.(2) and Drozd et al.(3) that discuss lifetimes of SCI for bimolecular reaction and unimolecular decomposition.

A: We thank the reviewer for pointing us to these references, which we now include in the revised paper. Besides, we have clarified in the Sect. 3.2 that considering that a portion of SCIs might undergo unimolecular decomposition, the SCIs yield estimated here represented a lower limit. In the Sect. 3.3 the results demonstrated that the OH yield was not obviously affected by RH, we speculated that the fraction of SCIs that decomposed and formed OH radicals was small.

Q14. Line 265-267 Specify the lifetimes for each loss process. You need to define "relatively long lifetime" with current quantitative estimates and relate this to unimolecular decomposition.

A: Thanks for your suggestion and we have defined the lifetimes of SCIs in the Sect. 3.2. In previous studies, the unimolecular decomposition of SCIs and the reactions of SCIs with water showed strong structure dependence. For CH_2OO and anti-CH_3CHOO, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were less than 1 ms, while their lifetimes of unimolecular reaction were much longer (Lin et al., 2016; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2012). For syn-CH_3CHOO and $(CH_3)_2COO$, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were more than 100 ms, while their lifetimes of unimolecular reaction were just few milliseconds at 298 K (Drozd et al., 2017; Huang et al., 2015; Long et al., 2018; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2014). For larger SCIs, their atmospheric lifetimes of unimolecular reaction and bimolecular reactions with H_2O and $(H_2O)_2$ were not quantified. Tillmann et al. (2010) inferred that about 46% ECIs formed from α-pinene ozonolysis would stabilize and at 44% RH about 65% SCIs formed H_2O_2 after reacting with water and 28% SCIs decomposed, Yao et al. (2014) showed that a large fraction of ECIs formed from α-cedrene were stabilized and the unimolecular decomposition of SCIs was suppressed by their bimolecular reactions. Although we could not estimate the rates of SCIs decomposition and their reaction with water, the results here proved that reaction with water was an essential route for limonene SCIs, and the rapid decomposition of HAHPs made an important contribution to H_2O_2 formation. At high RH, the unimolecular decomposition of limonene SCIs could be suppressed by their bimolecular reactions with water monomer and water dimer.

Q15. Line 310 What reactions were suppressed? Support or contrast this result with literature on SCI reactivity.

A: We regret for the unclear expression and we have better explained this in the revised manuscript. These suppressed reaction channels include the unimolecular decomposition of SCIs and the reactions of SCIs with other products in the system. For SCIs containing three or less carbon atoms, their unimolecular decomposition and reactions with water showed strong structure dependence, which have been clarified in the revised text. However, as for large SCIs their reaction pathways are still not clear. Tillmann et al. (2010) inferred that about 46% ECIs formed from α-pinene ozonolysis would stabilize and at 44% RH about 65% SCIs formed H_2O_2 after reacting with water and 28% SCIs decomposed, Yao et al. (2014) showed that a large fraction of ECIs formed from α-cedrene were stabilized and the unimolecular decomposition of SCIs was suppressed by their bimolecular reactions. The results in this study indicated that at high RH the unimolecular decomposition of limonene SCIs could be suppressed by their bimolecular reactions with water monomer and water dimer.

Q16. Line 317. You need to add equations that show how you calculate these yields, possibly in the supplemental.

A: We have added equations in the Supplement.

Q17. Section 3.5.3 You need to show in an equation how you calculate the relative amounts of stable and unstable peroxides.

A: We have added the equation in Sect. 3.5.3.

Q18. Line 479 Why was a MW of 300 g/mol assumed? Support with references or appropriate justification.

A: The average molecular weight of peroxides in particles is assumed to be 300 g/mol, which is estimated to be slightly less than the molecular weight of peroxyhemiacetals, and this value has been used to calculate the mass of particulate peroxides in some studies (Docherty et al., 2005; Nguyen et al., 2010; Surratt et al., 2006). We clarify this in the revised manuscript.

Q19. Line 504 Awkward sentence. What does “chemical effect” mean in this case?

A: We are sorry for not explaining that clearly. The chemical effect means the effect of water participating in some gas-phase and particle-phase reactions resulting in more low-volatility peroxides formation. This is now better explained in the revised text.

Q20. There are many grammatical errors and awkward sentences. Below is a partial list of these. Check over the entire manuscript again.

13 – “radical pha chemistry” 25 “Considerable generation of H2O2 from SOA in the aqueous se” 28 “SOA composition” (and throughout paper) 39-40 “suggest that the. . .need further study” 45 “Total monoterpene emissions are estimated to be..” 47 “non-negligible. . .” 64 “esters” 79 “reactions of alkenes. . .” 103 “Due to abundant water vapor...” 169 “first generation” 222 “constituents were generated” 280 “prior to generating

aldehyde” 293 “There was little to no effect of scavenger” 385 “chamber studies” 516“HACE might be generated. . .”

A: Thanks for your suggestion and we have revised all of these mistakes mentioned above. We also check over the manuscript seriously and correct other errors.

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Ahmad, W., Coeur, C., Cuisset, A., Coddeville, P., and Tomas, A.: Effects of scavengers of Criegee intermediates and OH radicals on the formation of secondary organic aerosol in the ozonolysis of limonene, J. Aerosol. Sci., 110, 70–83, doi: 10.1016/j.jaerosci.2017.05.010, 2017.

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~~Understanding the oxidants transition and SOA property~~The oxidation regime and SOA composition in limonene ozonolysis: ~~Role~~Roles of different double bonds, radicals~~radical chemistry~~, and water

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Abstract. Volatile organic compounds (VOCs) play an important role in air quality and climate change, largely because of their contribution to the formation of oxidizing compounds and secondary organic aerosol (SOA). ~~atmospheric oxidation capacity and secondary organic aerosol (SOA) formation through their oxidation~~. In this study, a series of products including peroxides and carbonyl compounds in both gaseous and particulate phases were simultaneously detected to help us investigate the ~~oxidants transition~~oxidation regime and SOA ~~property~~composition in limonene ozonolysis. ~~Reactants ratio, OH radical scavenger and relative humidity (RH) were controlled to discuss the effect of endocyclic and exocyclic double bonds (DBs), radial chemistry and water.~~ The roles of different double bonds (DBs), radicals, and water were also discussed. For the first issue, we were focused on the representative oxidizing compounds produced in limonene ozonolysis including stabilized Criegee intermediates (SCIs), OH radicals, and peroxides. ~~Alkene ozonolysis not only consumed but also regenerated oxidants, which made a great impact on atmospheric chemical processes. For this issue, we first paid attention to the generation of stabilized Criegee intermediates (SCIs) and OH radical.~~ The ~~variation~~dependence of H_2O_2 and hydroxymethyl hydroperoxide (HMHP) formation ~~with~~on RH ~~showed~~demonstrated that the reaction with water was an important reaction pathway for limonene SCIs, ~~the~~

~~importance of the reaction with water for limonene SCIs,~~ and the ~~estimated~~lower-limit SCIs yields of endocyclic and exocyclic DBs were estimated to be ~ 0.24 and ~ 0.43, respectively. OH yield was determined by adding sufficient OH scavenger, and the OH yields of endocyclic and exocyclic DBs were ~ 0.65 and ~ 0.24, respectively. The results indicated that in limonene ozonolysis the endocyclic DB was inclined to generate OH ~~radical~~radicals through hydroperoxide channel, while the exocyclic DB had higher fraction of forming SCIs. Besides, other gas-phase and particle-phase peroxides were also studied. The formation of peroxyformic acid (PFA) and peroxyacetic acid (PAA) were promoted significantly by the increasing RH and the oxidation degree, and the discrepancy between the experimental and model results suggested some missing formation pathways. Considerable generation of H_2O_2 ~~generation~~ from SOA in aqueous phase was observed especially at high $[O_3]$/[limonene], which was mainly attributed to the hydration and decomposition of ~~particulate~~ unstable peroxides in SOA such as peroxycarboxylic acids and peroxyhemiacetals. ~~Different DBs and radical chemistry revealed their influence on aerosol property through affecting the behavior of SOA on generating H_2O_2. As a species owning high SOA formation potential, another key issue we investigated in limonene ozonolysis was SOA property, for which peroxides and carbonyls were chosen as representatives.~~Different DBs and OH scavengers were found to have a great impact on the particulate peroxides, whose stabilities indicated that the types of peroxides in SOA were changed under different conditions. As for the contribution of peroxides to SOA, ~~The results showed that in limonene SOA,~~the results demonstrated that peroxides could account for ~~0.07–0.19~~less than 0.2 in limonene SOA at low $[O_3]$/[limonene], while~~ and 0.40–0.58~~ at high $[O_3]$/[limonene] peroxides could account for 0.4–0.6 in SOA.~~, which confirmed the important contribution of peroxides to aerosol formation.~~ The partitioning behavior of peroxides showed that ~~multigeneration~~multi-generation oxidation helped produce more low-volatility peroxides, which provided some explanation for higher SOA yield. The partitioning behavior of carbonyls was also discussed and the experimental partitioning coefficients (K_p)

were usually several orders of magnitude higher than theoretical values.~~, yet the relationship of K_p observed in laboratory with vapor pressure offered some reference for predicting the contribution of carbonyls to SOA formation.~~ This study provided new insights into the oxidation regime and SOA composition ~~oxidants transition and SOA property~~ in limonene ozonolysis, and limonene showed its specificity in many aspects when both endocyclic and exocyclic DBs were ozonated. We suggested that the atmospheric implications of terpenes containing more than one DB and the ~~properties of particulate products especially~~ SOA composition especially particulate peroxides ~~still needed~~need further study.

1 Introduction

As an important monoterpene, limonene has a high emission rate both from biogenic and anthropogenic sources, which is only second to pinene (Atkinson and Arey, 2003; Clausen et al., 2001; Fellin and Otson, 1994; Griffin et al., 1999; Guenther et al., 1995; Lamb et al., 1993; Seifert et al., 1989; Sindelarova et al., 2014; Wolkoff et al., 2000). Total monoterpene ~~emission amount was~~emissions are estimated to be 50 Tg C yr^{-1}, and limonene might comprise about 20% of that (Stroud et al., 2005). In addition to the ~~large quantity of emission~~massive emissions from vegetation, its extensive utilization in household and industrial processes also makes limonene ~~unnegligible~~ non-negligible in the atmosphere. Compared with its isomer α-pinene and β-pinene, an obvious feature of limonene in structure is that it owns two different double bonds (DBs): an endocyclic one and an exocyclic one, which makes the behavior and the fate of limonene in atmosphere complicated. Ozonolysis is an important reaction pathway of limonene and it serves as a source of free radicals, intermediate products, and aerosol. The first step of alkene ozonolysis is the addition of O_3 to the carbon-carbon double bond forming an energy-rich primary ozonide (POZ), which decomposes to two sets of carbonyls plus carbonyl oxides called excited Criegee intermediates (ECIs) (Criegee, 1975; Fenske et al., 2000; Gutbrod et al., 1996, 1997; Kroll et al., 2001a, b). ECIs can isomerize through hydroperoxide channel followed by OH production, rearrange to esters with subsequent decomposition, or undergo collisional stabilization forming stabilized Criegee intermediates (SCIs) (Cremer et al., 1993; Presto and Donahue, 2004; Richard et al., 1999; Wegener et al., 2007).

In the last few years, the oxidizing compounds formed in alkene ozonolysis have received much attention because they own the power of removing a series of trace gases and contribute much to atmospheric oxidation capacity (Möller, 2009; Prinn, 2003; Taatjes et al., 2013). The representatives of these compounds include OH radicals, SCIs, peroxides, etc. Investigating the oxidation regime of alkene ozonolysis is essential to understand the formation of these critical products and evaluate the actual impact of this reaction in the atmosphere. As an important source of OH radicals to make OH-initiated reactions continue in the dark (Kroll et al., 2001a, b), OH formation pathways in alkene ozonolysis have been extensively studied and the major pathway is considered to be the unimolecular decomposition of ECIs. Besides, some studies suggested that SCIs could also generate OH radicals through self-decomposition or reaction with other species (Anglada et al., 2002; Hasson et al., 2003; Kroll et al. 2001a, b; Tillmann et al., 2010; Zhang and Zhang, 2005). SCIs generated during the process prove to have sufficient lifetime, which is usually on the order of sub-seconds (Kroll et al., 2001a, b; Long et al., 2018; Mauldin et al., 2012), to react with other trace species, such as H_2O, SO_2, NO_2, carbonyls, alcohols, carboxylic acids, etc. (Sakamoto et al., 2017; Sipilä et al. 2014; Yao et al., 2014). There have been numbers of studies focusing on the reactions of SCIs containing 3 or less carbon atoms, and their unimolecular decomposition and reactions with water show strong structure dependence. The reactions with water monomer and water dimer are the main reaction pathways for CH_2OO and *anti*-CH_3CHOO, while for *syn*-CH_3CHOO and $(CH_3)_2COO$ the unimolecular decomposition would be more important. As for larger SCIs formed from biogenic alkenes ozonolysis, their reaction mechanisms are still unclear since it is difficult to synthesis and observe them directly. Furthermore, reactions of alkenes with ozone can also generate considerable amount of peroxides, which are considered to be important due to their oxidizability and role as radicals reservoir. H_2O_2 is the most crucial oxidant to oxidize S (IV) forming sulfuric acid and sulfate in the aqueous phase (Calvert et al., 1985; Penkett et al., 1979; Peña et al., 2001). And as a species accounting for 40–50% of the total global organic peroxides (Crounse et al., 2006; Khan et al., 2015), peroxycarboxylic acids (RC(O)OOH) play an important role in promoting atmospheric oxidation capacity and enhancing the acidity of aqueous phase.

~~Terpene oxidation is a well-known source of secondary organic aerosol (SOA), and limonene has proven to have higher potential of SOA formation than α-pinene because it is doubly unsaturated. The considerable potential of producing low-volatility compounds makes limonene have an important contribution to SOA formation (Andersson-Sköld and Simpson 2001; Kroll and Seinfeld, 2008; Lane et al., 2008; Lee et al., 2006). Some studies have investigated the initial process and SOA yield of limonene oxidation (Grosjean et al., 1992, 1993; Glasius et al., 2000; Leungsakul et al., 2005; Pathak et al., 2012), however, the knowledge of detailed reaction mechanism and SOA property in limonene ozonolysis remains unclear, especially when it comes to the effect of different DBs, which deserves efforts of further study to help us obtain a better understanding of limonene chemistry and its implications in the atmosphere.~~

~~Reactions of alkene with ozone have been explored by numerous researches because of their role as important sources of free radicals, intermediate products, and aerosol. The first step of alkene ozonolysis is the addition of O_3 to the carbon-carbon double bond forming an energy-rich primary ozonide (POZ), which decomposes to two sets of carbonyls plus carbonyl oxides called excited Criegee intermediates (ECIs) (Criegee, 1975; Fenske et al., 2000; Gutbrod et al., 1996, 1997; Kroll et al., 2001a, b). ECIs can isomerize through hydroperoxide channel followed by OH production, rearrange to eaters with subsequent decomposition, or undergo collisional stabilization forming stabilized Criegee intermediates (SCIs) (Cremer et al., 1993; Presto and Donahue, 2004; Richard et al., 1999; Wegener et al., 2007). Some products formed from alkene ozonolysis that own the power~~

~~of removing oxidizable compounds play an extremely important role in atmospheric chemical processes and evolution. It means that alkene ozonolysis is not only a process of consuming oxidants, but also a process of regenerating oxidants. These species contribute much to the atmospheric oxidation capacity and control the atmospheric self-cleaning process by removing a series of trace gases (Möller, 2009; Prinn, 2003; Taatjes et al., 2013). In recent years the reactive oxygen species (ROS), which includes oxygen-related free radicals (e.g., OH,~~ ~~HO_2, and RO_2), ions and molecules (e.g., H_2O_2, organic and inorganic peroxides), attracts increasing attention because of its relation with atmospheric oxidative level and adverse health effect caused by particle-phase ROS (Gallimore et al., 2017; Huang et al., 2016; Wragg et al., 2016). Alkene ozonolysis is thought to be an important source of OH radical in the atmosphere, which makes OH-initiated reactions possible to continue in the dark (Kroll et al., 2001a, b). SCIs generated during the process prove to have sufficient lifetime to react with other trace species, such as H_2O, SO_2, NO_2, carbonyls, alcohols, carboxylic acids, etc., enhancing the atmospheric oxidation capacity and promoting the SOA formation (Sakamoto et al., 2017; Sipilä et al. 2014; Yao et al., 2014). Furthermore, reactions of alkene with ozone can also generate considerable amount of peroxides, which receive much attention due to their oxidizability and role as radicals reservoir. H_2O_2 is the most crucial oxidant to oxidize S (IV) forming sulfuric acid and sulfate in the aqueous phase, and organic peroxides could also oxidize some species (Calvert et al., 1985; Penkett et al., 1979; Peña et al., 2001). Overall, although alkene ozonolysis consumes ozone, these critical compounds formed from the reaction contribute much to atmospheric oxidizing capacity. The existing knowledge of oxidants transition in alkene ozonolysis is not sufficient, yet it is essential for us to evaluate the actual impact of this reaction on atmospheric evolution and human health.~~

One important reason for limonene chemistry drawing much attention is its high SOA formation potential, which is proven to be higher than that of α-pinene because limonene is doubly unsaturated (Andersson-Sköld and Simpson 2001; Kroll and Seinfeld, 2008; Lane et al., 2008; Lee et al., 2006). Although progress has been made over the past years on simulating SOA formation with the theory of gas-particle partitioning, there are still large discrepancies between the model and experimental results (Cocker et al., 2001; Griffin et al., 1999; Hoffmann et al., 1997; Odum et al., 1996; Pankow, 1994; Presto et al., 2005; Pye and Seinfeld, 2010). Laboratory studies about SOA formation in limonene ozonolysis mainly focused on the aerosol yields under different conditions and identifying some products in the ~~particulate~~particle phase (Calogirou et al., 1999; Glasius et al., 2000; Grosjean et al., 1992, 1993; Leungsakul et al., 2005; Ng et al., 2006; Pathak et al., 2012), however, the composition ~~and property~~ of limonene SOA still need detailed study especially when it comes to the effect of different DBs. As a double-unsaturated terpene, SOA formation process of limonene could be more complicated than single-unsaturated terpene, as the ~~multigeneration~~ multi-generation oxidation has significant influence on SOA. In this study, two classes of species: peroxides and carbonyls are chosen to study their behaviors in limonene SOA formation. In the last few years, organic peroxides have been analyzed and suggested to be an important composition in aerosol (Docherty et al., 2005; Heaton et al., 2007; Li et al., 2016; Mertes et al., 2012; Pathak et al., 2012), and particulate peroxides could cause negative health effect after penetrating into lungs (Verma et al., 2009; Wragg et al., 2016). The reactive uptake and particle-phase reactions of carbonyls are believed to be responsible for fractions of aqueous SOA formation (Ervens et al., 2011; Mcneill et al., 2012), especially for dicarbonyls glyoxal and methylglyoxal. Up to now, few researches pay attention to the contribution of peroxides and carbonyls to limonene SOA and their behaviors in aerosol formation are vague in our knowledge.

~~Other key issues we are concerned about are the water and radical chemistry effect. Due to the sufficient water vapor in the atmosphere, it makes a great difference on several chemical and physical processes. It is necessary to design and conduct experiments in a wide range of relative humidity (RH) to provide more valuable information when the results obtained in laboratory are applied to the actual circumstances. Researches before~~

~~have provided evidence for OH radical formation in ozonolysis experiments and OH scavenger is often used to avoid the disturbance of OH reaction. However, it should be noted that when OH scavenger removes OH radical from the reactor, it could bring other reactions into the system, which would affect the production of hydroperoxy (HO_2) and alkylperoxy (RO_2) radicals. Some studies suggested that the choice of OH scavenger influenced the HO_2/RO_2 and SOA yield (Docherty and Ziemann, 2010; Keywood et al., 2004), yet its effect on other products was seldom discussed. It was proposed that to simulate the real atmospheric environment, experiments should be conducted at high HO_2/RO_2 and low reactant concentration (Jonsson et al., 2008a). Chew and Atkinson (1996) argued that the ability of 2-butanol and cyclohexane in scavenging OH radical was similar, however, 2-butanol was proven to give more HO_2 radical than cyclohexane, suggesting that 2-butanol was a more appropriate scavenger. Here, to study the impact of water and radical chemistry on the reaction system, a series of experiments were preformed under various RH conditions in the presence or absence of different OH scavengers (2-butanol or cyclohexane).~~

The focus of this study is to investigate the ~~oxidants transition and SOA property~~ oxidation regime and SOA composition in limonene ozonolysis, especially for the roles~~role~~ of different DBs, ~~radical chemistry~~radicals, and water. On the one hand, through determining the generation of SCIs, OH ~~radical~~radicals, and peroxides, the ~~issue of oxidants transition~~ oxidation regime in the reaction system is discussed. On the other hand, peroxides and carbonyls are taken as ~~representative~~representatives to study their behaviors in SOA formation. ~~Reactants ratio, OH scavenger, and RH are controlled to explore the effect of different DBs, radical chemistry, and water.~~

2 Experimental

2.1 Chemicals

R-(+)-Limonene (Sigma-Aldrich, ⩾ 99.0%), 2-butanol (Sigma-Aldrich, 99.5%), cyclohexane (Alfa Aesar, ⩾ 99.9%), potassium iodide (KI, Alfa Aesar, 99.9%), hydrogen peroxide (H_2O_2, Alfa Aesar, 35wt.%), ortho-phosphoric acid (H_3PO_4, Fluka, 85–90%), hemin (Sigma, ⩾ 98.0%), 4-hydroxyphenylacetic acid (Alfa Aesar, 99%), formaldehyde (Sigma-Aldrich, 37wt.%), acetaldehyde (Amethyst Chemicals, 40wt.%), acetone (Fluka, ⩾ 99.7%), hydroxyacetone (Sigma-Aldrich, 90%), glyoxal (Sigma-Aldrich, 40wt.%), methylglyoxal (Sigma-Aldrich, 40wt.%), 2-butanone (Alfa Aesar, ⩾ 99%), acetonitrile (Alfa Aesar, ⩾ 99.7%), 2,4-dinitrophenyl hydrazinel (DNPH, TCI, 50wt.%), ammonia solution ($NH_3{\cdot}H_2O$, Beijing Tongguang Fine Chemicals Company, 25.0–28.0%), ammonium chloride (NH_4Cl, Beijing Tongguang Fine Chemicals Company, ⩾ 99.5%), sulfuric acid (H_2SO_4, Xilong Chemical Company, 95.0–98.0%), ultrapure water (18MΩ, Millipore), N_2 (⩾ 99.999%, Beijing Haikeyuanchuang Practical Gas Company Limited, Beijing, China), O_2 (⩾ 99.999%, Beijing Haikeyuanchuang Practical Gas Company Limited, Beijing, China), polytetrafluoroethylene (PTFE) filter membrane (Whatman Inc., 47mm in diameter), and quartz microfiber filters (Whatman Inc.) were used in this study.

2.2 Apparatus and procedures

A flow tube reactor (2 m length, 70 mm inner diameter, quartz wall) equipped with a water jacket for controlling temperature was used to investigate the ozonolysis of limonene. All the experiments were conducted at 298±0.5 K in the dark. O_3 was generated by O_2 photolysis in a 2 L quartz tube with a low-pressure Hg lamp, and the detailed quantification method of O_3 was described in our previous study (Chen et al., 2008). H_2O_2 produced by UV irradiation of O_2 and trace water was measured in control experiments and deducted from the results. Limonene gas was generated by passing N_2 flow over liquid limonene in a diffusion tube at the selected temperature and OH scavenger (2-butanol or cyclohexane) gas was generated with a bubbler. The concentrations of limonene and OH scavenger were determined by gas chromatography with flame ionization detector (GC-FID, Agilent 7890A, USA). Water vapor was generated by passing N_2 through a water bubbler, which contained a carborundum disc submerging in ultrapure water (18 MΩ). The mixing gas including limonene, OH scavenger, ozone, and dry or wet synthetic air (80% N_2 and 20% O_2), was successively introduced into the reactor, and with a total flow rate of 2 standard L min^{-1}, the residence time was estimated to be 240 s.

To explore the reaction mechanism of endocyclic and exocyclic DBs ozonolysis, and the effect of ~~multigeneration~~ multi-generation oxidation in limonene ozonolysis, two sets of experiments with different ratios of ozone to limonene concentration were conducted. In the following content, [O_3] denoted the concentration of ozone, [limonene] denoted the concentration of limonene, and [O_3]/[limonene] denoted the ratio of ozone to limonene concentration. In the low [O_3]/[limonene] set of experiments, the initial concentrations of limonene and ozone were ~ 280 ppbv and ~ 500 ppbv, respectively. In the high [O_3]/[limonene] set of experiments, the initial concentrations of limonene and ozone were ~ 183 ppbv and ~ 19 ppmv, respectively. In both sets of experiments, enough 2-butanol and cyclohexane were added to scavenge OH ~~radical~~ radicals in the RH range of 0–90%. In the tables and figures, the low and high ratio sets of experiments were denoted with ~~mark~~marks L and H, and the conditions in the absence of scavenger, in the presence of 2-butanol and in the presence of cyclohexane were represented by No-sca, 2-But, and C-hex, respectively. Experimental conditions were listed in Table 1.

According to previous studies, in limonene ozonolysis the rate constant of endocyclic DB reaction with ozone was 2×10^{-16} cm^3 $molecule^{-1}$ s^{-1} (Atkinson, 1990; Shu and Atkinson, 1994), while the exocyclic DB reaction with ozone was about 30 times slower than endocyclic DB (Zhang et al., 2006). Based on those rate constants we estimated that at low [O_3]/[limonene], less than 1% exocyclic DB was ozonated, so this situation mainly represented the ~~firstgeneration~~ first-generation oxidation. In this circumstance, because the ozone concentration was low, OH reaction would impact the amount of limonene consumed by O_3. In the presence of OH scavenger, ~ 42% endocyclic DB reacted with O_3, while in the absence of scavenger, ~ 38% endocyclic DB reacted with O_3. At high [O_3]/[limonene], more than 99% endocyclic DB and about 51% exocyclic DB reacted with ozone, and since the ozone concentration in this situation was high, the OH effect on ozonolysis was presumed to be unimportant. The latter condition, which contained ~~multigeneration~~ multi-generation oxidation process, was more likely to happen since the ratio of [O_3] to [limonene] was similar to the ratio in the real atmosphere.

It should be noted that one advantage of flow tube reactor was that the wall would be in equilibrium with the ~~gaseous~~ gas phase after a stationary period, and according to our observation, this process usually needed about 2 h. In order to stabilize the system and diminish the wall effect as much as possible, the reactor was usually aged for 2 h prior to measurement and after experiments the reactor was rinsed out with ultrapure water and blown to dry with N_2.

2.3 Products analysis

To better investigate the gas-particle partitioning of products formed in limonene ozonolysis, we analyzed gas-phase and particle-phase products simultaneously. ~~The formation of total~~ Total peroxides and a series of low-molecule-weight (LMW) peroxides were measured here, and in the following discussion we regarded the peroxides that could be detected by high performance liquid chromatography (HPLC) as LMW peroxides, while considered the peroxides undetermined by HPLC as high-molecule-weight (HMW) peroxides. For particle-phase peroxides detection, a PTFE filter was used for SOA collection and the mass of SOA was measured by semi-micro balance (Sartorius, Germany). Since the control experiment results showed that long-time collection led to the loss of some peroxides in particles, the collection time was controlled to be 3 h for each filter, and the accuracy of particulate products analysis was discussed in the Supplement. Each loaded PTFE filter was extracted with 20 mL H_3PO_4 solution (pH 3.5) using a shaker (Shanghai Zhicheng ZWY 103D, China) at 180 rpm and 4 °C for 15 min~~, then the SOA solution was analyzed at once. The~~ and the extraction efficiency was confirmed in our previous work (Li et al., 2016).~~, and this method could be regarded as a reliable way to determine particulate peroxides.~~ ~~The peroxides that were detected by high performance liquid chromatography (HPLC) were regarded as LMW peroxides, while for the peroxides undetermined by HPLC, we considered them as high-molecule weight (HMW) peroxides in the following discussion. In SOA extract solution, LMW peroxides were hardly detected, indicating that the particle-phase total peroxides concentration could be treated as the particle-phase HMW peroxides concentration. For gas-phase peroxides detection, gas through the filter was collected in a coil collector with H_3PO_4 solution (pH 3.5), which was detected immediately.~~ For gas-phase peroxides detection, the air samples through the filter were collected in a glass coil collector at a temperature of 4 °C with H_3PO_4 solution (pH 3.5) serving as the rinsing solution. The coil collector is around 30 cm long and its effective length is about 100 cm. The coil is similar with that used in earlier studies (Grossmann et al., 2003; Sauer et al., 1996, 1997) and the diagram of the thermostatic coil collector is shown in the Supplement.

The detection method of peroxides was reported in our previous studies (Hua et al., 2008; Li et al., 2016), so only a brief description was given here. LMW peroxides were analyzed by HPLC (Agilent 1100, USA) coupled with post-column derivatization and fluorescence detection on line. Peroxides separated by column chromatography reacted with *p*-hydroxyphenylacetic acid (POPHA) under the catalysis of hemin forming POPHA dimers, and then the dimers were quantified by fluorescence detector. The standard solution of peroxides was prepared and used for calibration in each measurement. The detection limit of gas-phase LMW peroxides was about 22 pptv, and in the particle phase the detection limit was around 0.068 ng/μg for H_2O_2. The accuracy of this method was estimated to be around 7% and the precision of the measurement results was usually within 20%. The concentration of total peroxides (H_2O_2, ROOH, and ROOR′) was determined by iodometric spectrophotometric method, which based on the reaction of peroxides and iodide ions (Docherty et al., 2005; Mutzel et al., 2013). Briefly, excessive KI solution was added into samples purged of O_2, after staying 12–24 h in the dark for derivatization, the I_3^- ions produced were quantified at 420 nm by UV/VIS spectrophotometer (SHIMADZU UV-1800, Japan). The efficiency of total peroxides measurement was discussed in our previous work (Li et al., 2016) and the detection limit of total peroxides was about 0.92 ppbv in the gas phase and 0.025 μg/μg in the particle phase. The accuracy of this method was estimated to be 10% and the precision of the measurement results was usually within 20%.

To measure particle-phase carbonyls, SOA was collected onto a quartz microfiber filter for 3 h, after collection the filter was put upside down in a conical flask. 5 mL acetonitrile, 1 mL DNPH saturated solution, and 50 μL H_2SO_4 solution (0.25 M) were added into the flask in sequence, the flask was then shaken at 180 rpm and 4 °C for 3 h and kept in dark for 12–24 h waiting for detection afterwards. For gas-phase carbonyls measurement, gas through the filter was directly introduced into a Horibe tube, which was placed in a cold trap (Beijing Tiandijingyi TH-95-15-G, China) at about -98 °C to freeze the products in tube. The Horibe tube ~~was~~ is made of an inlet tube (25 cm length, 4 cm O.D.), a coil (7 laps, 1 cm O.D.), and an outlet tube with a carborundum disc. The diagram of the Horibe tube is shown in the Supplement. After collection, 10 mL acetonitrile was added to rinse the Horibe tube inside to dissolve carbonyl compounds, and then the solution was mixed with DNPH saturated solution and H_2SO_4 solution to derivatize for 12–24 h in the dark. Samples derived were analyzed by HPLC with UV detection (Agilent 1100, USA), and details of the process could be found in our previous work

(Wang et al., 2009). The detection limit of carbonyl compounds was about 0.3 ppbv in the gas phase, and in the particle phase it was 0.15 ng/μg in terms of formaldehyde (FA). The accuracy of this method was estimated to be 8% and the precision of the measurement results was within 30%. For acetone (ACE) and methylglyoxal (MGL), since the yields of both were low their uncertainty might be high during the analysis.

2.4 Wall loss experiments

To make results more accurate, we designed and conducted a series of control experiments to quantify the wall loss effect~~. Two types of experiments were carried out~~, including gaseous products and aerosol wall loss evaluation.~~ As for gaseous products, peroxides and carbonyls were chosen to analyze.~~ Gas containing peroxide constituents was generated by passing N_2 through a diffusion tube controlled at 4 °C, which contained certain peroxide solution in it. The synthetic method of multiple organic peroxides was described in our previous work (Huang et al., 2013). Gas containing carbonyl constituents was prepared by injecting liquid substance into an evacuated steel canister (15 L, Entech Instrument), and then N_2 was added constantly until the pressure in canister reached 30 psi. The outlet of canister was linked with a mass flow controller to regulate the gas flow rate. The gas containing peroxides or carbonyls was mixed with synthetic air at different RH and introduced into the reactor with a rate of 2 standard L min^{-1}, and the concentrations of peroxides and carbonyls were controlled to be at the similar level with the products observed in limonene ozonolysis. After the gas mixture was introduced into the flow tube, around 2 h was needed for gas and wall to become balanced, and then the measurement would start. The aerosol wall loss experiment used two-stage reaction equipment containing two flow tube reactors. Because we wanted to explore the wall loss effect on pure SOA, the first reactor was used to generate aerosol where limonene and O_3 had sufficient time to react completely. The gas containing aerosol out of the first reactor was mixed with synthetic air at different RH and introduced into the second reactor with a rate of 2 standard L min^{-1}. The particles at the inlet and outlet of the second reactor were collected on PTFE filters and measured by balance to calculate the SOA concentration. The wall loss fraction of gas-phase constituent or SOA was determined as the difference between the inlet and outlet ~~concentration~~ concentrations divided by the inlet concentration as Eq. (S1) (in the Supplement). ~~, which could be expressed as ([In] – [Out]) / [In].~~ The wall loss experiments were conducted in the RH range of 0–90%, and the profiles of loss fractions as a function of RH could be used for correcting products and SOA yields to diminish the wall loss effect.

3 Results and discussion

3.1 Wall loss correction

~~The wall loss fractions of four kinds of LMW peroxides observed in limonene ozonolysis were discussed here, i.e., H_2O_2, hydroxymethyl hydroperoxide (HMHP), peroxyformic acid (PFA), and peroxyacetic acid (PAA). Figure 1 showed the dependence of these peroxides wall loss fractions on RH. The four loss profiles indicated increasing~~ The wall loss fractions of LMW peroxides increased with increasing RH, and this tendency was especially obvious for ~~PFA~~peroxyformic acid (PFA) and ~~PAA~~peroxyacetic acid (PAA), whose loss fractions increased successively with RH. For H_2O_2 and ~~HMHP~~hydroxymethyl hydroperoxide (HMHP), their wall loss fractions went up quickly above 50% RH, yet didn't have large change below 50% RH. Generally speaking, HMHP had the highest wall loss fraction, which could reach ~ 0.25 at 90% RH, while PFA had the lowest wall loss fraction. The average of these LMW peroxides loss fractions was used to correct the wall loss effect for gas-phase HMW peroxides.

~~The wall loss effect on a series of carbonyls formed in limonene ozonolysis was evaluated in the RH range of 0–90%, and the profiles as a function of RH were shown in Fig. 2. In general, the relationship of carbonyls~~ As for carbonyl compounds, the variation of their wall loss fractions with RH was not very obvious. The loss curves of ~~formaldehyde (FA)~~FA, acetaldehyde (AA), and ~~acetone (ACE)~~ ACE were kind of irregular. For hydroxyacetone (HACE), glyoxal (GL), and ~~methylglyoxal (MGL)~~MGL, their wall loss fractions were lowest at 10% RH, and then they gradually arose with increasing water vapor concentration. The higher loss fractions of GL and MGL compared with other carbonyls could be attributed to their inclination towards hydration. FA had the lowest wall loss fraction around 0.03, while MGL had the highest loss fraction around 0.12.

The wall loss effect on SOA was also discussed from 0% to 90% RH. The SOA wall loss fraction was ~ 0.06 at dry condition, then increased slightly with water vapor concentration. At 50% RH the SOA loss fraction was ~ 0.11, and at 90% RH, the SOA loss fraction was ~ 0.17. ~~Detailed information was~~ The dependencies of peroxides, carbonyls, and SOA wall loss fractions on RH were definitely shown in the Supplement.

3.2 SCIs generation

~~The reaction channels of ECIs are complex in monoterpene ozonolysis, and some studies suggested that the unimolecular decomposition and intermolecular stabilization were dominant pathways (Aschmann et al., 2002; Chew and Atkinson, 1996; Lin et al., 2014; Ma et al., 2008; Tillmann et al., 2010). Due to the relatively long lifetime of SCIs, bimolecular reactions of SCIs with other trace species are possible. The reaction between SCIs and water has received much attention because sufficient water vapor exists in the atmosphere, thus it is regarded as an important reaction pathway for SCIs. This class of reaction, which produces α-hydroxyalkyl hydroperoxides (HAHPs) decomposing to H_2O_2, carbonyls, and carboxylic acids, is thought to be an essential source of these compounds and serve as a principle source of H_2O_2 formation without light (Becker et al., 1990; Becker et al., 1993; Gäb et al., 1985; Sauer et al., 1999). However, because of the complicated structure and the difficulty of synthesis, it is not easy to observe monoterpene SCIs reaction directly, resulting in that the reaction~~

~~mechanism and the rate constant of monoterpene SCIs reaction with water are still unclear. Vereecken et al. (2017) reported that the concentration of biogenic SCIs in the atmosphere was strongly limited by unimolecular decay, yet the obvious formation of products from HAHPs decomposition under humid condition in monoterpene ozonolysis demonstrated that SCIs reaction with water was important (Anglada et al., 2002; Ma et al., 2008; Tillmann et al., 2010). Jiang et al. (2013) suggested that the formation of HAHPs was the most favorable pathway for limonene SCIs reaction with H_2O, and the subsequent decomposition of HAHPs was thought to be prior to generate aldehyde and H_2O_2 (Chen, 2016; Kumar, 2014). Although theoretical calculation results indicated that HAHPs decomposition was slow, some studies proved that water and acid molecule could greatly promote the decomposition process (Anglada et al., 2002; Anglada et al., 2011; Aplincourt and Anglada, 2003; Crehuet et al., 2001), and the H_2O_2 formation from HAHPs decomposition was very fast (Chen et al., 2016; Winterhalter et al., 2000). The fact that few HAHPs larger than HMHP were identified in alkene ozonolysis also provided evidence for the rapid decomposition of large HAHPs. In this study, through investigating hydroperoxides formation from 0% to 90% RH in limonene ozonolysis, we tried to provide more information about limonene SCIs generation and their reaction with water.~~

The reaction between SCIs and water, which produces α-hydroxyalkyl hydroperoxides (HAHPs) decomposing to H_2O_2, carbonyls, and carboxylic acids, is thought to be an essential source of these compounds and serve as a principle source of H_2O_2 formation without light (Becker et al., 1990, 1993; Gäb et al., 1985; Sauer et al., 1999). Figure 1~~3 showed~~ shows the dependence of H_2O_2 yield and HMHP yield on RH, and the six profiles in each subgraph ~~represented~~ represents conditions at low or high $[O_3]$/[limonene] in the presence or absence of OH scavenger (2-butanol or cyclohexane). The molar yield used here ~~was~~ is defined as the ratio of products molar number to the molar number of limonene consumed as Eq. (S2) (in the Supplement). It was obvious that although in different cases, the variation of H_2O_2 yield and HMHP yield with RH had similar tendency, both of which increased significantly from 0% to 70% RH, and then they approached the limiting values. ~~The effect of OH scavenger was not obvious.~~ There was little effect of OH scavenger. At low [O_3]/[limonene], the maximum H_2O_2 yield was ~ 24.00% without OH scavenger, ~ 24.60% with 2-butanol, and ~ 22.95% with cyclohexane, respectively. At high [O_3]/[limonene], the maximum H_2O_2 yield reached ~ 41.20% without OH scavenger, ~ 41.80% with 2-butanol, and ~ 40.50% with cyclohexane. As for HMHP, its yield was much higher at high [O_3]/[limonene] (~ 5.43%) than at low [O_3]/[limonene] (~ 0.62%), and the specific information could be found in the following picture. It is usually believed that H_2O_2 has two generation pathways, one is HO_2 self-reaction, and the other one is HAHPs decomposition. The former pathway is considered as a main contributor to H_2O_2 during daytime, while the latter is regarded as a route without photochemistry. In the reaction system discussed here, when we applied a box model coupled with limonene reaction mechanism extracted from the Master Chemical Mechanism (MCM) v3.3 (website: http://mcm.leeds.ac.uk/MCMv3.3.1) to simulate the reaction, it was estimated that the yield of H_2O_2 formed from HO_2 self-reaction was less than 0.1% under both dry and wet conditions. When 2-butanol or cyclohexane reaction mechanism was taken into consideration the contribution of this pathway to H_2O_2 formation was still very limited, hence, it was assumed that HO_2 self-reaction was not important for H_2O_2 generation in this reaction system. Under dry condition, a small amount of water vapor might desorb from the flow tube wall and participated in reactions resulting in a little H_2O_2 formation. Thus the high H_2O_2 yield observed here might be mainly attributed to SCIs reaction with water.

From 0% to 70% RH, both of H_2O_2 yield and HMHP yield were promoted significantly by the increasing RH, indicating that the reaction with water gradually turned to be dominant for limonene SCIs. Jiang et al. (2013) suggested that the formation of HAHPs was the most favorable pathway for limonene SCIs reaction with H_2O, and the subsequent decomposition of HAHPs was thought to be prior to generating aldehyde and H_2O_2 (Chen, 2016; Kumar, 2014). Although theoretical calculation results indicated that HAHPs decomposition was slow, some studies proved that water and acid molecule could greatly promote the decomposition process (Anglada et al., 2002; Anglada et al., 2011; Aplincourt and Anglada, 2003; Crehuet et al., 2001), and the H_2O_2 formation from HAHPs decomposition was very fast (Chen et al., 2016; Winterhalter et al., 2000). The fact that few HAHPs larger than HMHP were identified in alkene ozonolysis also provided evidence for the rapid decomposition of large HAHPs. Above 70% RH, the appearance of the limiting values of H_2O_2 yield and HMHP yield suggested that the water vapor concentration was high enough to make the bimolecular reaction of SCIs and H_2O suppress~~ other reaction channels.~~ the unimolecular decomposition of SCIs and the reactions of SCIs with other products. In previous studies, the unimolecular decomposition of SCIs and the reactions of SCIs with water showed strong structure dependence. For CH_2OO and *anti*-CH_3CHOO, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were less than 1 ms, while their lifetimes of unimolecular reaction were much longer (Lin et al., 2016; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2012). For *syn*-CH_3CHOO and $(CH_3)_2COO$, the atmospheric lifetimes of their bimolecular reactions with H_2O and $(H_2O)_2$ were more than 100 ms, while their lifetimes of unimolecular reaction were just few milliseconds at 298 K (Drozd et al., 2017; Huang et al., 2015; Long et al., 2018; Sheps et al., 2014; Taatjes et al., 2013; Welz et al., 2014). For larger SCIs their atmospheric lifetimes of unimolecular reaction and bimolecular reactions with H_2O and $(H_2O)_2$ were not quantified. Tillmann et al. (2010) inferred that about 46% ECIs formed from α-pinene ozonolysis would stabilize and at 44% RH about 65% SCIs formed H_2O_2 after reacting with water and 28% SCIs decomposed, Yao et al. (2014) showed that a large fraction of ECIs formed from α-cedrene were stabilized and the unimolecular decomposition of SCIs was suppressed by their bimolecular reactions. Although we could not estimate the rates of SCIs decomposition and their reaction with water, ~~The~~the results here proved that reaction with water was an essential route for limonene SCIs, and the rapid decomposition of HAHPs made an important contribution to H_2O_2 formation. At high RH, the unimolecular decomposition of limonene SCIs could be suppressed by their bimolecular reactions with water monomer and water dimer.

According to experimental conditions elaborated in Sect. 2.2, we tried to calculate the contribution of endocyclic DB and exocyclic DB ozonolysis to H_2O_2 and HMHP formation, and furthermore, infer the SCIs generation in different DBs ozonolysis. The predicted SCIs yield was derived by combining the limiting yields of H_2O_2 and HMHP together as Eq. (S3) (in the Supplement), based on the assumption that the limiting yield of H_2O_2 was equal with the large SCIs yield (Hasson, 2001a, b). Considering that a portion of SCIs might undergo unimolecular decomposition a lower-limit yield of SCIs formed in limonene ozonolysis was estimated here. Additionally, in Sect. 3.3 the results demonstrated that the OH yield was not obviously affected by RH, we speculated that the fraction of SCIs that decomposed and formed OH radicals was small.~~The SCIs yield estimated here could be regarded as a lower bound as a small fraction of SCIs might undergo decomposition.~~ It was observed that OH scavenger didn't have a huge impact on the SCIs measurement results, while big difference existed between the two DBs ozonolysis. The SCIs yield of endocyclic DB ozonolysis was around 24.45%, yet the exocyclic DB ozonolysis had larger stabilization fraction of ECIs, which was about 42.90%. It should be noted that since our experiments just isolated the ozonolysis of different DBs according to their discrepancy in reaction rates, the SCIs yield of exocyclic DB ozonolysis estimated here was not for a specific product formed from endocyclic DB ozonolysis, but the sum of first-generation products with a remaining double bond, as shown in the Fig. 2. ~~It~~ The results meant that even though exocyclic DB ozonolysis was much slower than endocyclic DB, it played ~~an unnegligible~~a non-negligible role in generating SCIs in limonene ozonolysis.

3.3 OH ~~radical~~radicals generation

~~In the last few years, OH formation pathways in alkene ozonolysis were extensively studied, and the major~~

~~pathway was considered to be the unimolecular decomposition of ECIs. Besides, some studies suggested that SCIs could also generate OH radical through self-decomposition or reaction with other species (Anglada et al., 2002; Hasson et al., 2003; Kroll et al. 2001a, b; Tillmann et al., 2010; Zhang and Zhang, 2005). When it came to OH formation in alkene ozonolysis, one thing still under debate was the effect of water because of the large discrepancy among the existing publications. Anglada et al. (2002) indicated that water could increase OH production using quantum mechanical calculations. Nevertheless, observed influence of RH on OH production was limited. Tillmann et al. (2010) reported higher OH yield under humid condition, but more studies showed that OH formation was independent of water vapor concentration (Aschmann et al., 2002; Atkinson and Aschmann, 1993; Atkinson et al., 1992; Berndt et al., 2003; Forester and Wells, 2011; Hasson et al., 2003). OH yield in alkene ozonolysis is highly dependent on the reactant molecular structure, and as far as we know, only Herrmann et al. (2010) researched OH yield of both DBs in limonene ozonolysis under dry condition. OH radical could be directly detected by laser-induced fluorescence (LIF), or be indirectly determined using OH radical scavenger. 2-butanol and cyclohexane are both commonly used OH scavengers, and OH yield can be determined by detecting the amount of 2-butanone generated from 2-butanol reaction with OH or the amount of cyclohexanone plus cyclohexanol formed from cyclohexane reaction with OH. Aschmann et al. (2002) suggested that using 2-butanol to measure OH formation would be more accurate than using cyclohexane.~~

In this study, OH yield was determined by adding sufficient 2-butanol as scavenger and measuring how much 2-butanone generated. Aschmann et al. (2002) suggested that using 2-butanol to measure OH formation would be more accurate than using cyclohexane. The yield of 2-butanone formed from the reaction of 2-butanol and OH had been detected previously (Aschmann et al., 2002; Baxley and Wells, 1998; Chew and Atkinson, 1996), and we used 0.66 as an average of the reported values to calculate OH ~~radical~~radicals yield as the same as Forester and Wells (2011). The results showed that the yield of OH ~~radical~~radicals produced from endocyclic

DB ozonolysis was 0.65 ± 0.21, while for exocyclic DB, OH yield was 0.24 ± 0.13. OH generation didn't show evident dependence on water vapor concentration, and the OH yield of endocyclic DB ozonolysis was in the range of values published, while the OH yield of exocyclic DB ozonolysis was determined to be higher than that Herrmann et al. (2010) reported. The effect of water on OH formation in alkene ozonolysis is still under debate because of the discrepancies among the existing publications. Anglada et al. (2002) indicated that water could increase OH production using quantum mechanical calculations and Tillmann et al. (2010) reported higher OH yield under humid condition. Nevertheless, more studies showed that OH formation was independent of water vapor concentration (Aschmann et al., 2002; Atkinson and Aschmann, 1993; Atkinson et al., 1992; Berndt et al., 2003; Forester and Wells, 2011; Hasson et al., 2003). The fact that OH yields of both unsaturated bonds were not obviously affected by RH suggested that the major pathway of OH generation in limonene ozonolysis was decomposition of ECIs through hydroperoxide channel. The possibility that other OH formation pathways also existed could not be totally excluded, but we speculated that the contribution of other pathways to OH formation was not significant. Hence it could be concluded that in limonene ozonolysis, the endocyclic DB was inclined to generate OH ~~radical~~radicals through ECIs decomposition, while the exocyclic DB had higher fraction of stabilization forming SCIs.

3.4 Peroxycarboxylic acids generation

~~As a species accounting for 40–50% of the total global organic peroxides (Crounse et al., 2006; Khan et al., 2015), peroxycarboxylic acids (RC(O)OOH) play an important role in promoting atmospheric oxidation capacity and enhancing the acidity of aqueous phase. Here the generation of PFA (CH(O)OOH) and PAA ($CH_3C(O)OOH$) were observed in limonene ozonolysis. These two kinds of peroxycarboxylic acids got growing attention recent years as reactive oxidants, and PFA had shown to be more active than PAA. In field~~

~~observations, PAA was reported widely in remote and urban areas (Lee et al., 2000; Liang et al., 2013; Zhang et al., 2010), yet there were few reports about PFA existence in the atmosphere (Liang et al., 2015), which might be attributed to the rapid decomposition of PFA or its precursor. The dominant formation pathway of PAA was considered as the reaction of $CH_3C(O)OO$ radical with HO_2 (Groβ et al., 2014; Lightfoot et al., 1992; Winiberg et al., 2016), while the formation pathway of PFA hadn't been identified yet and was speculated to be the reaction of HC(O)OO radical with HO_2 (Liang et al., 2015).~~ In our experiments, two kinds of peroxycarboxylic acids PFA (CH(O)OOH) and PAA ($CH_3C(O)OOH$) were observed, ~~In this study~~ and the yields of PFA and PAA ~~produced in limonene ozonolysis~~ both showed large ~~discrepancy~~ discrepancies between the low and high ratio sets of experiments (Fig. ~~4~~3). The yields of PFA and PAA at high $[O_3]$/[limonene] were about three times as much as those at low $[O_3]$/[limonene]. When the exocyclic DB was ozonated, the generation of PFA and PAA were enhanced to a large extent compared with only endocyclic DB ozonolysis, indicating that exocyclic DB ozonolysis had an important impact on PFA and PAA. As far as we know, this was the first time that such high yields of PFA and PAA in alkene ozonolysis were reported. For both of PFA and PAA, the highest molar yields were observed when no OH scavenger was used, demonstrating that OH reaction contributed to part of their formation. When OH ~~radical was~~radicals were scavenged, we speculated that PFA and PAA formation were mainly through ECIs isomerization and decomposition following HC(O)OO and $CH_3C(O)OO$ radicals generation. These radicals could further react with HO_2 forming peroxycarboxylic acids, which provided a plausible explanation for higher yields of PFA and PAA observed in experiments with 2-butanol than with cyclohexane.

The yields of PFA and PAA had positive relationship with water vapor concentration in all the cases, reaching the highest level at 90% RH. Model simulation results significantly underestimated the PAA formation and ignored the RH effect, indicating some missing pathways of PAA generation related with water. Because of the deficiency of PFA mechanism in MCM we didn't simulate the PFA formation, however, ~~the positive correlation between the yields of PFA and PAA~~ the similar variation tendency of PFA and PAA yields could provide some evidence for the assumption that they might have similar forming mechanism. In previous studies, the dominant formation pathway of PAA was considered as the reaction of $CH_3C(O)OO$ radicals with HO_2 radicals (Groβ et al., 2014; Lightfoot et al., 1992; Winiberg et al., 2016), while the formation pathway of PFA hadn't been identified yet and was speculated to be the reaction of HC(O)OO radicals with HO_2 radicals (Liang et al., 2015). The rate constant of RO_2 reaction with HO_2 had been investigated in a series of studies and was thought to be unaffected by water (Atkinson et al., 1999; Lightfoot et al., 1992; Tyndall et al., 2001; Wallington et al., 1992). Therefore the increase of PFA and PAA yields with RH might be attributed to the promoting effect of water on HC(O)OO and $CH_3C(O)OO$ radicals generation. Because the highly reactive peroxides were easy to lose on the wall, ~~studies in~~ chamber studies, which usually last for hours, might be hard to observe these peroxycarboxylic acids. In field observations, PAA are reported widely in remote and urban areas (Lee et al., 2000; Liang et al., 2013; Zhang et al., 2010), yet there are few reports about PFA existence in the atmosphere (Liang et al., 2015). The results here demonstrated that limonene ozonolysis could contribute to PFA and PAA formation, and although the high instability and reactivity made PFA difficult to observe, it might have a short stay in the atmosphere.

3.5 Particulate peroxides and H_2O_2 generation

3.5.1 H_2O_2 evolution at low $[O_3]$/[limonene]

H_2O_2 generation from SOA in aqueous phase ~~was~~ is thought to be a possible way of producing H_2O_2 in cloud water or releasing H_2O_2 continuously after inhalation. Here we provided the quantitative measurement of H_2O_2 generation from SOA produced in limonene ozonolysis in different cases. Considering that when SOA extract solution was kept at room temperature (298 K) some unstable constituents would decay rapidly, we chose to keep the SOA solution at 4 °C to maintain the stability of samples and prolong their storage time. ~~resulting in that the evolution process couldn't be observed completely, we chose to keep the SOA solution at 277 K for several days in the dark to determine the change of total peroxides concentration and H_2O_2 concentration successively.~~ In the low ratio set of experiments, the concentration of total peroxides in solution nearly maintained stable in 48 h and H_2O_2 concentration could reach a steady state after going through a short rising period. After aqueous H_2O_2 concentration reached plateau, the amount of particulate H_2O_2 per particle mass formed in different cases from 0% to 90% RH was calculated and shown in Table 2. No matter OH scavenger was used or not, SOA produced at higher RH was inclined to own higher capacity of producing H_2O_2. When no OH scavenger was used, the lowest H_2O_2 content per particle mass was ~ 1.13 ng/μg at dry condition, and the highest value was ~ 2.45 ng/μg at 90% RH. When 2-butanol was used the changing trend of H_2O_2 generation resembled the condition without scavenger, and the minimum was ~ 1.33 ng/μg at 0% RH, while the maximum was ~ 2.89 ng/μg at 80% RH, which was a little higher than that at 90% RH. It was interesting to note that, in the presence of cyclohexane, the trend of H_2O_2 generation in SOA solution differed greatly from both of the above. Even under dry condition, the H_2O_2 content per particle mass could reach ~ 3.22 ng/μg, and the maximum was ~ 4.63 ng/μg at 90% RH.

3.5.2 H_2O_2 evolution at high $[O_3]$/[limonene]

In the experiments of high $[O_3]$/[limonene], SOA produced with different scavengers was found to have different rates of generating H_2O_2 in solution at ~~277 K~~4 °C, according to which we determined appropriate detection frequency and total duration for the three kinds of SOA. SOA produced without OH scavenger had the lowest rate of generating H_2O_2, so an eight-day measurement result was reported here. For SOA produced in the presence of 2-butanol, it was a little faster than the former on generating H_2O_2 and reached the limiting value within six days. However, SOA produced with cyclohexane had a much faster rate than both of the above, and the H_2O_2 concentration in solution became stable within three days. For all of the three kinds of SOA, the total peroxides concentration decreased slightly in the analysis duration, which was concretely clarified in the Supplement.

It ~~was~~is obvious that SOA produced in the high ratio set of experiments ~~had~~has greater ability of generating H_2O_2, and Fig. 4~~5~~ ~~showed~~shows the time profiles of H_2O_2 evolution of different kinds of SOA. When no OH scavenger was used, H_2O_2 concentration in solution rose constantly in about six days then became stable. The limiting value of H_2O_2 generation was influenced by RH, which was ~ 5.28 ng/μg at 0% RH, then increased gradually with increasing RH until 80% RH (~ 14.45 ng/μg). In the experiments with 2-butanol, H_2O_2 concentration kept rising in the first four days then became stable. At dry condition, the limiting value of H_2O_2 content was ~ 7.60 ng/μg, then it increased until 50% RH. SOA produced above 50% RH didn't show obvious difference, and the highest H_2O_2 content was ~ 15.50 ng/μg. When cyclohexane was added, H_2O_2 concentration in solution rose up quickly in the first day then tended to be stable. The limiting value of H_2O_2 generation was also affected by water vapor concentration, yet the promoting effect was not very significant and no big difference was observed above 30% RH. The limiting H_2O_2 content per particle mass was ~ 16.64 ng/μg at dry condition and was ~ 30.00 ng/μg above 30% RH.

3.5.3 Different stabilities of particulate peroxides

According to the measurement results of total peroxides and H_2O_2 in SOA solution, the particulate peroxides could be roughly divided into two categories: one ~~was~~is unstable that could decompose or hydrate to generate H_2O_2, the other one ~~was~~is stable that could nearly remain unchanged in several days. Based on the assumption that all the peroxides contained one peroxy bond, the molar fractions of peroxides with different stabilities under various conditions could be calculated as follows:

$$\varphi_i = \frac{N_i}{N_{total}} \tag{1}$$

Where φ_i is the molar fraction of stable or unstable peroxides in particles, N_i (mol) is the molar number of stable or unstable peroxides in particles, and N_{total} (mol) is the molar number of total peroxides in particles.~~.~~ ~~The molar fraction used here represented the molar number of stable or unstable particulate peroxides to the molar number of total peroxides in SOA.~~ At low [O_3]/[limonene], the molar fraction of ~~unstable particulate peroxides~~ unstable peroxides in SOA was around 0.11–0.13 in the case of adding no scavenger, which was similar with the case of adding 2-butanol. In the presence of cyclohexane, the molar fraction of unstable ~~particulate~~ peroxides would increase to 0.20–0.32. At high [O_3]/[limonene], the molar fractions of unstable peroxides in the case of adding no scavenger and adding 2-butanol were also similar, both of them ranged in 0.13–0.25, yet this value would reach ~ 0.50 when cyclohexane was used. Detailed information was shown in the Supplement. Model results showed that most peroxides produced in the reaction were ROOH and R(O)OOH, in addition, some studies proposed that peroxyhemiacetals also made important contribution to aerosol (Tobias and Ziemann, 2000; Tobias et al., 2000). The acetal reaction producing peroxyhemiacetals was reversible, so part of the peroxyhemiacetals in SOA were possible to hydrolyze and form some peroxides. To investigate the stabilities of ROOH and R(O)OOH, we synthesized methyl hydroperoxide (MHP) and ethyl hydroperoxide (EHP) to represent ROOH. As for R(O)OOH, PFA and PAA were used as representative. All the synthesis solutions were stored at ~~277 K~~ 4°C, which was the same with the experimental condition. PFA and PAA were found to decompose and generate H_2O_2 in several days, yet MHP and EHP maintained stable. Hence we speculated that peroxycarboxylic acids and peroxyhemiacetals might be the main components of unstable

~~particulate~~ peroxides in particles and contribute to H_2O_2 generation in aqueous phase.

~~SOA formed with different OH scavengers had different molar fractions of unstable peroxides, which indirectly proved the influence of radical chemistry on SOA composition. When cyclohexane was used as OH scavenger, the formation of more unstable species might be attributed to the extra RO_2 radicals provided by cyclohexane that participated in subsequent reactions.~~ The amount of H_2O_2 generation at low $[O_3]$/[limonene] measured here was comparable with the published value of H_2O_2 produced from α-pinene SOA in solution (Li et al., 2016; Wang et al., 2011). However, at high $[O_3]$/[limonene], H_2O_2 generation level increased significantly, which proved that the ~~multigeneration~~multi-generation oxidation improved the formation of peroxycarboxylic acids and peroxyhemiacetals in particles. The results demonstrated that SOA produced in limonene ozonolysis could behave as an important source of H_2O_2 in aqueous phase, and they also showed the difference between the SOA formed from single-unsaturated monoterpene ozonolysis and double-unsaturated monoterpene ozonolysis.

3.5.4 The effect of OH scavenger

Researches before have provided evidence for OH radicals formation in ozonolysis experiments and OH scavenger is often used to avoid the disturbance of OH reaction. 2-butanol and cyclohexane used here are both commonly used OH scavengers, and Chew and Atkinson (1996) showed that there was no difference in their abilities to scavenge OH radicals. However, it should be noted that OH scavenger could convert OH radicals into a mixture of hydroperoxy (HO_2) and alkylperoxy (RO_2) radicals and higher $[HO_2]/[RO_2]$ is observed when 2-butanol is used (Docherty and Ziemann, 2003; Jonsson et al., 2008a). As shown in the Sect. 3.6.1, the fact that both at low and high $[O_3]$/[limonene], the SOA yield was higher with 2-butanol than cyclohexane suggested that the increase of HO_2 concentration promoted the SOA formation. This result is consistent with that suggested by Keywood et al. (2004) who observed the higher $[HO_2]/[RO_2]$ resulting in higher SOA yield in cyclohexene ozonolysis, while Docherty and Ziemann (2003) showed that increased $[HO_2]/[RO_2]$ inhibited aerosol formation in β-pinene ozonolysis.

Some studies demonstrated that acid and peroxide products were sensitive to OH scavengers (Ahmad et al., 2017; Henry and Donahue, 2011; Ma et al., 2008). Here through investigating the stabilities of particulate peroxides formed in different conditions, we showed that the choice of OH scavenger would influence the types of particulate peroxides produced in the reaction. In the Sect. 3.5.3, we suggest that peroxycarboxylic acids and peroxyhemiacetals be the main components of unstable peroxides in particles. The peroxyhemiacetals formed by heterogeneous reactions of peroxides and aldehydes could dissociate into these species, yet the formation of peroxyhemiacetals was not supposed to be affected by OH scavenger since both of the peroxides and carbonyl compounds observed in the gas phase did not show large differences when OH scavenger changed. We speculate that the impact of OH scavenger on particulate peroxides could be mainly attributed to the formation of peroxycarboxylic acids under different $[HO_2]/[RO_2]$. Keywood et al. (2004) indicated that the OH scavenger impacted the HO_2 reaction with acylperoxy radicals forming acid and peracid products, which were considered to have low volatility. With higher HO_2 concentration, the SOA yield in the presence of 2-butanol was higher than that in the presence of cyclohexane, since the HO_2-acylperoxy reactions contributed to some low-volatility products. However, the RO_2 radicals formed from cyclohexane may also participate into reactions and help produce more unstable peroxycarboxylic acids which could partition into the particle phase.

3.6 Contribution of peroxides to SOA

3.6.1 SOA formation

The yield of SOA produced from limonene ozonolysis in different cases from 0% to 90% RH was measured. The SOA yield ~~was~~is defined as the ratio of aerosol mass concentration to the mass concentration of limonene consumed as Eq. (S4) (in the Supplement). In the six different cases, SOA yield was determined to be unaffected by water, yet it showed strong dependence on the reactants ratio and the use of OH scavenger, which was shown in Table 3. Whether at low or high [O_3]/[limonene], the case of none OH scavenger had the highest SOA yield, which suggested the effect of OH reaction on aerosol formation, and the order of SOA yield was no scavenger > 2-butanol > cyclohexane. The case owning the highest SOA yield was high [O_3]/[limonene] without OH scavenger (0.511 ± 0.097), while the lowest SOA yield was detected at low [O_3]/[limonene] with cyclohexane (0.288 ± 0.038).

3.6.2 Peroxides mass fraction

In this study, the iodometric method was used to analyze the particle-phase total peroxides content in limonene ozonolysis. The iodometric method ~~was~~is commonly regarded as a standard method of detecting total peroxides and it could almost quantify all kinds of peroxides (Bonn et al., 2004; Jenkin, 2004). The mass fraction used here ~~was~~is defined as the ratio of particulate peroxides mass to SOA mass as Eq. (S5) (in the Supplement). The average molecular weight of peroxides in particles ~~was~~is assumed to be 300 g/mol, which is estimated to be slightly less than the molecular weight of peroxyhemiacetals, and this value has been used to calculate the mass of particulate peroxides in some studies (Docherty et al., 2005; Nguyen et al., 2010; Surratt et al., 2006). ~~and the~~The mass fractions of peroxides in the six cases from 0% to 90% RH were summarized in Table 3. In each case, the peroxides mass fraction in SOA increased slightly with RH, and significant difference existed between the low and high ratio sets of experiments, while the effect of OH scavenger was little. The mass fraction of peroxides at [O_3]/[limonene] was usually below 0.2, however, at high [O_3]/[limonene] the peroxides mass fraction increased to 0.4–0.6. ~~. When no OH scavenger was used, the range of peroxides mass fraction was 0.065–0.169 at low [O_3]/[limonene], and 0.401–0.492 at high [O_3]/[limonene]. When using different scavengers, some changes were observed. In the presence of 2-butanol, peroxides mass fraction ranged in 0.101–0.189 at~~

~~low $[O_3]$/[limonene], and in 0.502–0.580 at high $[O_3]$/[limonene]. In the presence of cyclohexane, peroxides mass fraction was in the range of 0.087–0.189 at low $[O_3]$/[limonene], and 0.477–0.512 at high $[O_3]$/[limonene].~~ Docherty et al. (2005) reported that peroxides mass fraction in α-pinene SOA was ~ 0.47, while for β-pinene SOA the fraction was ~ 0.85. Li et al. (2016) observed that peroxides could account for ~ 0.21 in α-pinene SOA. Here we first reported the mass fraction of peroxides in SOA derived from limonene ozonolysis, highlighting the important role of organic peroxides in SOA composition, especially when ~~multigeneration~~multi-generation oxidation happened.

3.6.3 Peroxides partitioning

The molar yields of HMW peroxides in both ~~gas-phase~~gas phase and ~~particle-phase~~particle phase were determined in different cases. ~~The effect of RH and OH scavenger was not obvious,~~ There was little effect of RH and OH scavenger, while under high $[O_3]$/[limonene] the HMW peroxides yield increased by ~ 10% in contrast with the condition of low $[O_3]$/[limonene]. Furthermore, the gas-phase HMW peroxides and particle-phase HMW peroxides were discussed separately, which was shown in Fig. ~~6~~5. For gas-phase HMW peroxides, the molar yield showed a decreasing tendency with increasing RH, which was obvious at high $[O_3]$/[limonene]. For particle-phase HMW peroxides, the molar yield showed a slight increasing dependence on RH and their yield was promoted significantly by oxidizing degree. A possible explanation for the RH effect was that the water content in aerosol would increase when RH increased, and ~~promote~~promoted the uptake of some compounds into particles. However, some researches proved that water effect on the partitioning of organic products seemed to be small especially when no inorganic seed particles were used (Jonsson et al., 2006, 2008a, b), so we proposed that the promoting effect of water on particulate HMW peroxides could be mainly attributed to the ~~chemical~~ effect of water participating in some gas-phase and particle-phase reactions resulting in more low-volatility peroxides formation. The contribution of multigeneration-oxidation products to SOA formation in limonene ozonolysis had been stated by some studies (Hoffmann et al., 1997; Ng et al., 2006), and the results here indicated that organic peroxides might account for a considerable proportion of those products. On the one hand, ~~multigeneration~~multi-generation oxidation helped produce more low-volatility peroxides in the ~~gaseous~~gas phase that could partition into particles, and on the other hand, it could also accelerate the occurrence of some heterogeneous reactions through providing more reactive species.

3.7 Contribution of carbonyls to SOA

3.7.1 Carbonyls formation

Carbonyl compounds including HACE, FA, AA, ACE, GL, and MGL were detected in the reaction. Unlike peroxides, whose generation often showed obvious increasing dependence on RH, only FA yield would increase with RH and other carbonyls production were not significantly affected by water. The yields of HACE, AA, ACE, GL, and MGL in different cases were summarized in Table 4. The fact that HACE formed without OH scavenger but did not form when cyclohexane was used indicated that HACE might be generated ~~generate~~ from OH reactions. As for AA and ACE, both of their yields in the presence of cyclohexane were found to be lower than the case without scavenger, especially at low [O_3]/[limonene]. This suggested that OH reaction contributed to a potion of their formation and the endocyclic DB ozonolysis did not tend to generate AA and ACE. It was speculated that these three kinds of carbonyls could also generate from the reaction of 2-butanol or HO_2 radicals promoted their formation, so the presence of 2-butanol increased their yields. As regards GL and MGL, RH and OH scavenger did not make big influence, and both of their yields at high [O_3]/[limonene] were higher than the condition of low [O_3]/[limonene]. The generation of FA had a positive dependence on RH, which was specifically illustrated in the Supplement and the total variation range could be found in Table 4.

3.7.2 Experimental and theoretical partitioning coefficients

A parameter that has been used widely to describe the partitioning feature of a compound is described as follows (Pankow, 1994; Pankow and Bidleman, 1992):

$$K_{p,i} = \frac{F_i/TSP}{A_i}$$ (2~~1~~)

Where $K_{p,i}$ (m^3 μg^{-1}) is the partitioning coefficient of compound i, TSP (μg m^{-3}) is the concentration of total suspended particulate matter, F_i (μg m^{-3}) and A_i (μg m^{-3}) are the particulate and gaseous concentrations of compound i, respectively. The measured partitioning coefficients of FA, AA and ACE were on the magnitude of 10^{-5}, and the partitioning coefficients of HACE, GL, and MGL were on the magnitude of 10^{-4}.

Theoretical gas-particle partitioning coefficients of these compounds were calculated using the absorption equilibrium equation defined by Pankow (1994), which was used widely to estimate the ability of a substance to partition into the particulate phase and predict SOA yield in model (Griffin et al., 1999; Hohaus et al., 2015; Odum et al., 1996; Yu et al., 1999):

$$K_{p,i} = \frac{760\, R\, T\, f_{om}}{MW_{om}\, 10^6\, \zeta_i\, p_{L,i}^0}$$ (3~~2~~)

Where R is the ideal gas constant (8.206×10^{-5} m^3 atm mol^{-1} K^{-1}), T is temperature (K), f_{om} is the mass fraction of TSP that is the absorbing organic material (om), which is 1 here, MW_{om} (g mol^{-1}) is the mean molecular weight of om phase, which is estimated to be 130 g mol^{-1} in this study, ζ_i is the activity coefficient of compound i in the om phase, which is assumed to be unity, and $p_{L,i}^0$ (Torr) is the vapor pressure of compound i, which is predicted by the method of Moller et al. (2008). The calculated gas-particle partitioning coefficients of HACE, FA, AA, ACE, GL, and MGL were 3.972×10^{-8} m^3 μg^{-1}, 2.096×10^{-11} m^3 μg^{-1}, 1.624×10^{-10} m^3 μg^{-1}, 6.192×10^{10} m^3 μg^{-1}, 5.476×10^{-10} m^3 μg^{-1}, and 1.246×10^{-9} m^3 μg^{-1}, respectively, at 298K. The experimental K_p value of HACE was about 10,000 times bigger than the theoretical value, and for other carbonyls, the experimental K_p was about 100,000 times bigger than the theoretical value. The gap between the experimental K_p and predicted K_p of GL and MGL we estimated was comparable with the results of Healy et al. (2008, 2009), but higher than that of Ortiz et al. (2013). The fact that gas-particle partitioning coefficients of carbonyls observed were much higher than theoretical values indicated that carbonyl compounds made a more important contribution for SOA formation than estimated. Although the partitioning coefficients measured in experiments showed huge difference with calculated values, some relationship between the measured K_p and the vapor pressure of carbonyl compounds was observed. Figure 6~~7~~ showed the dependence of measured $lg(K_p)$ and predicted $lg(K_p)$ on $lg(p^0)$, and their linear fitting curves. The slope of the linear fitting equation of predicted $lg(K_p)$ versus $lg(p^0)$ was -0.964, and R^2 was 0.998. The slope of the linear fitting equation of measured $lg(K_p)$ versus $lg(p^0)$ was -0.484, and R^2 was 0.750, which indicated that the $lg(K_p)$ of carbonyls observed in laboratory also had negative correlation with $lg(p^0)$. A plausible explanation for the large difference between the measured and predicted K_p was that carbonyl compounds were easy to polymerize and react with other species on particles, resulting in that these carbonyls existed in forms of hydrates and oligomers in SOA (Corrigan et al., 2008; Hastings et al., 2005; Kroll et al., 2005; Volkamer et al., 2007). The hydrates and oligomers of carbonyls have much lower vapor pressures than their precursors, and they could reversibly return to their carbonyl monomers during analysis (Healy et al., 2008; Ortiz et al., 2013; Toda et al., 2014).~~and the relationship of measured K_p and p^0 might provide some reference for predicting the contribution of carbonyls to SOA formation.~~

4 Conclusions and implications

An experimental study about the ~~oxidants transition and SOA property~~ oxidation regime and SOA composition in limonene ozonolysis with respect to the roles~~role~~ of different DBs, ~~radical chemistry~~radicals, and water was reported in this work. ~~To investigate the oxidants transition in this reaction system, a~~A series of products owning the oxidizing capacity including SCIs, OH ~~radical~~radicals, and peroxides in both gaseous and particulate phases were detected. Based on the variation of H_2O_2 and HMHP generation on RH, the importance of limonene SCIs reaction with water was confirmed and the yield of SCIs could be estimated, which was ~ 0.24 for endocyclic DB and ~ 0.43 for exocyclic DB. OH ~~radical~~ yields of endocyclic and exocyclic DBs were indirectly determined to be ~ 0.65 and ~ 0.24, demonstrating the different reaction mechanisms of different DBs in limonene ozonolysis. The formation of two peroxycarboxylic acids PFA and PAA was observed and their high yields were first reported in alkene ozonolysis. The yields of PFA and PAA increased with RH and oxidizing degree, showing the effect of water and the exocyclic DB oxidation on their formation. The H_2O_2 generation from SOA in solution provided evidence for the ability of SOA to contribute oxidants in aqueous phase. And through investigating the stabilities of particulate peroxides formed in different conditions, we showed that the choice of OH scavenger would influence the types of particulate peroxides produced in the reaction.~~Particles produced at high [O_3]/[limonene] showed much higher potential of forming H_2O_2 than particles produced at low [O_3]/[limonene], and the difference in H_2O_2-generating rate and amount among particles formed with different OH scavengers demonstrated the impact of radical chemistry on SOA composition.~~ The partitioning behaviors of peroxides and carbonyls were discussed and the results showed their importance to SOA formation. Particulate peroxides could account for ~~0.07–0.19~~less than 0.2 in limonene SOA at low [O_3]/[limonene] and ~~0.40–0.58~~0.4–0.6 at high [O_3]/[limonene], which proved the important role of peroxides in SOA composition especially when ~~multigeneration~~ multi-generation oxidation happened. Due to their ability to polymerize and react with other species on particles, the carbonyl compounds often exist in SOA in forms of hydrates and oligomers whose vapor pressures are much lower than their precursors. The partitioning coefficients of carbonyls observed in laboratory were always several orders of magnitude higher than theoretical values, ~~since~~

~~their ability to polymerize and react with other species on particles~~ indicating that their contribution to SOA was higher than the estimation. ~~Through determining the formation of oxidizing species and the partitioning of peroxides and carbonyls, the issues of oxidants transition and SOA property in limonene ozonolysis were investigated.~~ Limonene ~~showed~~shows its specificity in many aspects because of its different DBs, suggesting that some influence of terpenes containing more than one DB in the atmosphere might be underestimated before. The results demonstrated the effect of oxidizing degree, radical chemistry, and water on limonene SOA ~~property~~composition, while ~~the structures and properties of~~ some particulate products especially peroxides should be studied further. The implications of limonene chemistry in the atmosphere were close with the ~~multigeneration~~multi-generation oxidation that happened on its different DBs, and whether the phenomena could be observed in other terpenes oxidation that contained more than one DB needed more laboratory evidence.

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Table 1. Experimental Conditions.

Exp.	[Limonene] (ppbv)	$[O_3]$ (ppbv)	OH Scavenger	[OH Scavenger] (ppmv)	RH (%)
L(No-sca)	280	500	—	—	0–90
L(2-But)	280	500	2-butanol	350	0–90
L(C-hex)	280	500	cyclohexane	420	0–90
H(No-sca)	183	19000	—	—	0–90
H(2-But)	183	19000	2-butanol	350	0–90
H(C-hex)	183	19000	cyclohexane	420	0–90

Note: L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane; RH, relative humidity.

Table 2. H_2O_2 generation per particle mass (ng/ug) in SOA formed with different OH scavengers in the relative humidity (RH) range of 0–90% under low $[O_3]$/[limonene] ratio.

	0% RH	10% RH	30% RH	50% RH	70% RH	80% RH	90% RH
No-sca	1.13±0.22	1.42±0.40	1.53±0.28	1.88±0.16	2.24±0.42	2.26±0.44	2.45±0.48
2-But	1.33±0.15	1.56±0.18	1.77±0.12	2.02±0.65	2.55±0.43	2.89±0.42	2.66±0.57
C-hex	3.22±0.52	3.95±0.43	4.12±0.40	4.22±0.33	4.63±0.24	4.26±0.33	4.63±0.96

Note: No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

Table 3. The SOA yield and mass fraction of particulate peroxides at low or high $[O_3]$/[limonene] ratio in the presence or absence of OH scavenger from 0% to 90% relative humidity (RH).

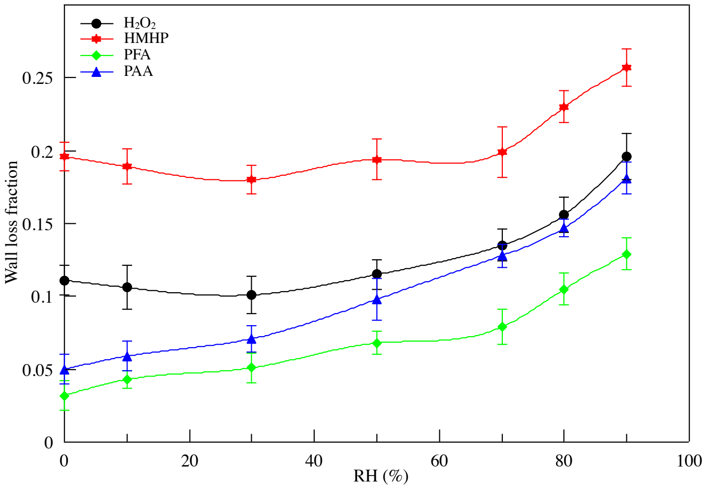
	L(No-sca)	L(2-But)	L(C-hex)	H(No-sca)	H(2-But)	H(C-hex)
0% RH	0.065±0.006	0.101±0.009	0.087±0.011	0.401±0.016	0.502±0.008	0.477±0.010
10% RH	0.091±0.010	0.124±0.013	0.113±0.009	0.436±0.009	0.534±0.009	0.502±0.013
30% RH	0.125±0.010	0.147±0.011	0.143±0.015	0.458±0.020	0.553±0.015	0.506±0.011
50% RH	0.149±0.007	0.174±0.011	0.161±0.016	0.466±0.016	0.571±0.009	0.512±0.007
70% RH	0.155±0.009	0.178±0.009	0.169±0.014	0.486±0.023	0.576±0.010	0.503±0.011
80% RH	0.169±0.013	0.189±0.010	0.189±0.012	0.492±0.015	0.580±0.013	0.502±0.014
90% RH	0.156±0.010	0.183±0.013	0.189±0.013	0.492±0.017	0.580±0.018	0.506±0.016
SOA Yield	0.379±0.039	0.337±0.048	0.288±0.038	0.511±0.097	0.479±0.044	0.401±0.068

Note: L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

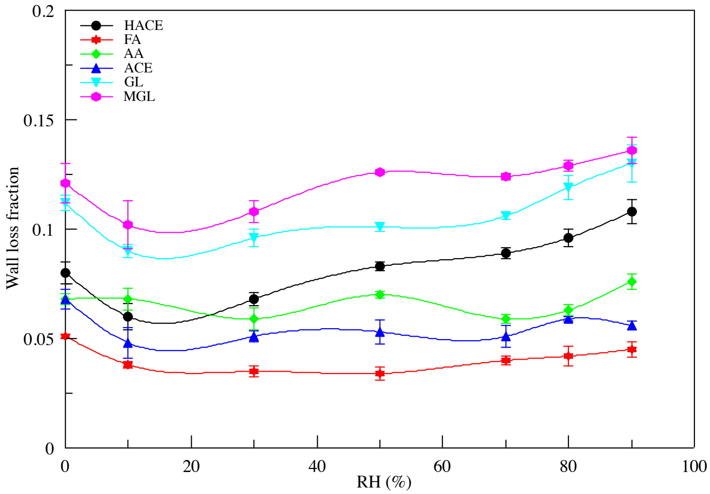
Table 4. Yields (%) of carbonyls at low or high $[O_3]$/[limonene] ratio in the presence or absence of OH scavenger.

	HACE	FA	AA	ACE	GL	MGL
L(No-sca)	2.04±0.48	7.02±0.90–10.58±0.94	1.32±0.24	0.22±0.15	0.89±0.25	0.56±0.34
L(2-But)	3.94±0.55	4.90±0.86–7.77±0.86	3.98±0.60	0.35±0.18	0.69±0.25	0.55±0.44
L(C-hex)	—	5.21±0.66–8.25±0.55	—	—	0.81±0.24	0.67±0.56
H(No-sca)	4.45±0.52	13.11±0.63–27.00±1.56	2.16±0.84	0.79±0.22	1.33±0.41	1.35±0.61
H(2-But)	10.15±2.11	11.03±0.77–23.33±0.62	7.86±1.32	1.00±0.28	1.25±0.36	1.31±0.21
H(C-hex)	—	10.80±1.28–23.32±1.21	1.89±1.22	0.60±0.47	1.25±0.51	1.23±0.22

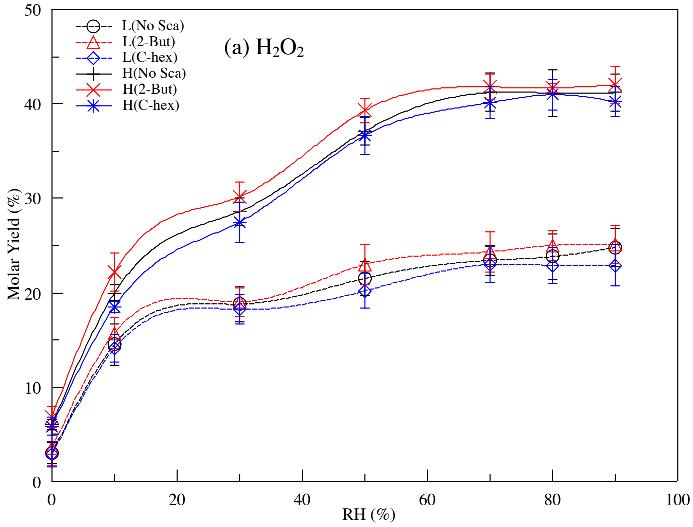
Note: —, below detection limit; L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane; HACE, hydroxyacetone; FA, formaldehyde; AA, acetaldehyde; ACE, acetone; GL, glyoxal; MGL, methylglyoxal.

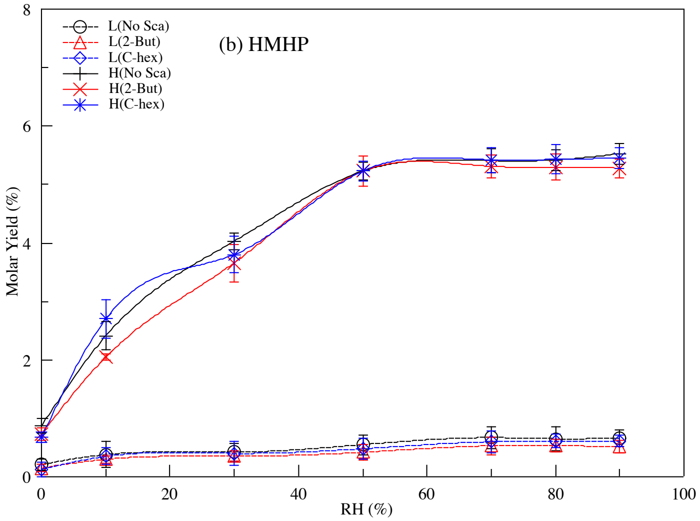


~~**Figure 1.** The dependence of peroxides wall loss fractions on relative humidity (RH). H_2O_2, hydrogen peroxide; HMHP, hydroxymethyl hydroperoxide; PFA, peroxyformic acid; PAA, peroxyacetic acid.~~



~~**Figure 2.** The dependence of carbonyls wall loss fractions on relative humidity (RH). HACE, hydroxyacetone; FA, formaldehyde; AA, acetaldehyde; ACE, acetone; GL, glyoxal; MGL, methylglyoxal.~~


(a) H_2O_2
L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
Molar Yield (%)
0
10
20
30
40
50
RH (%)
0
20
40
60
80
100


(b) HMHP
L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
Molar Yield (%)
0
2
4
6
8
RH (%)
0
20
40
60
80
100

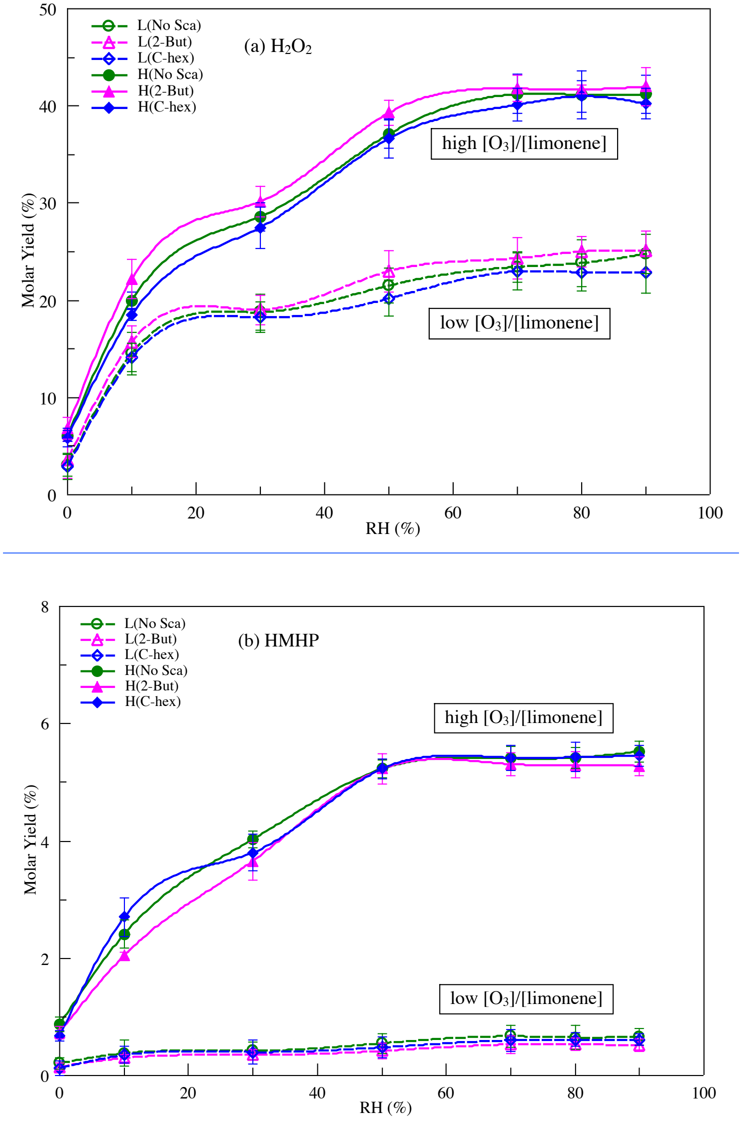


Figure ~~3~~1. Dependence of (a) H_2O_2 yield and (b) HMHP yield on relative humidity (RH) at low or high $[O_3]$/[limonene] ratio in the presence or absence of OH scavenger (2-butanol or cyclohexane). H_2O_2, hydrogen peroxide; HMHP, hydroxymethyl hydroperoxide; L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

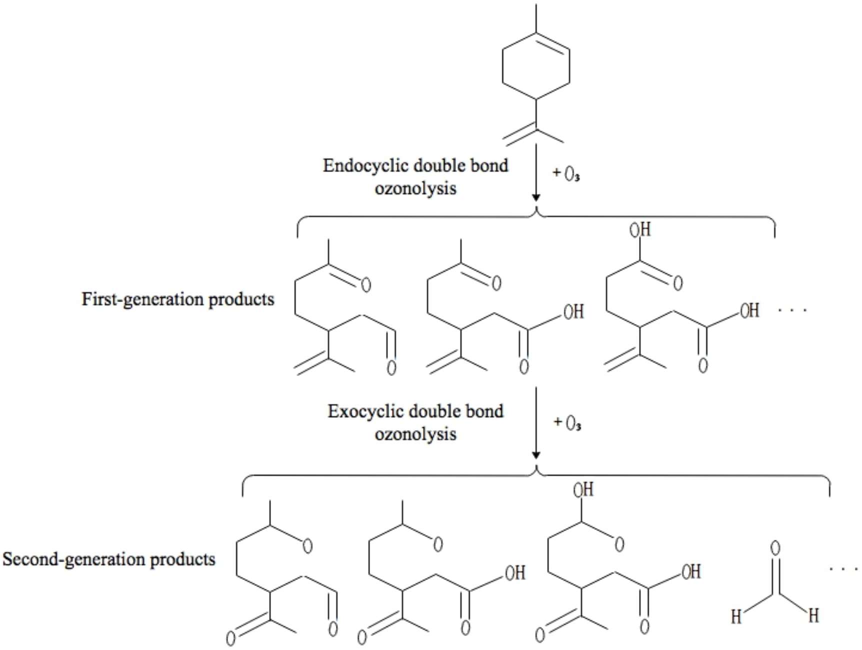
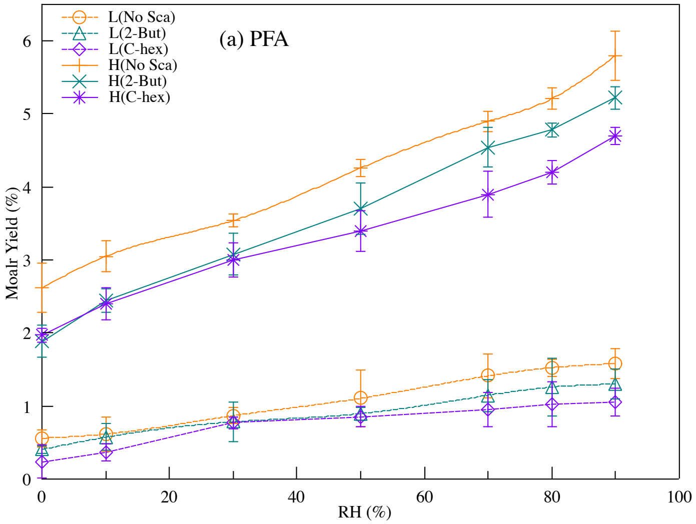
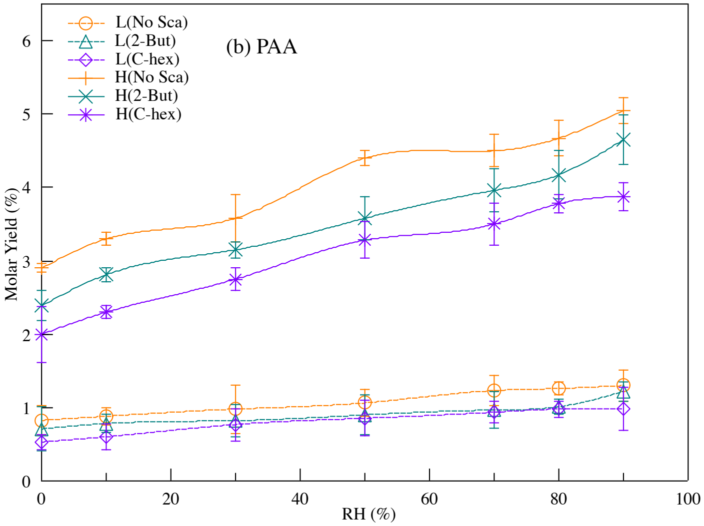


Figure 2. The reaction scheme of endocyclic and exocyclic double bonds in limonene ozonolysis


L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
(a) PFA
Moalr Yield (%)
0
1
2
3
4
5
6
0
20
40
60
80
100
RH (%)


L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
(b) PAA
Molar Yield (%)
0
1
2
3
4
5
6
0
20
40
60
80
100
RH (%)

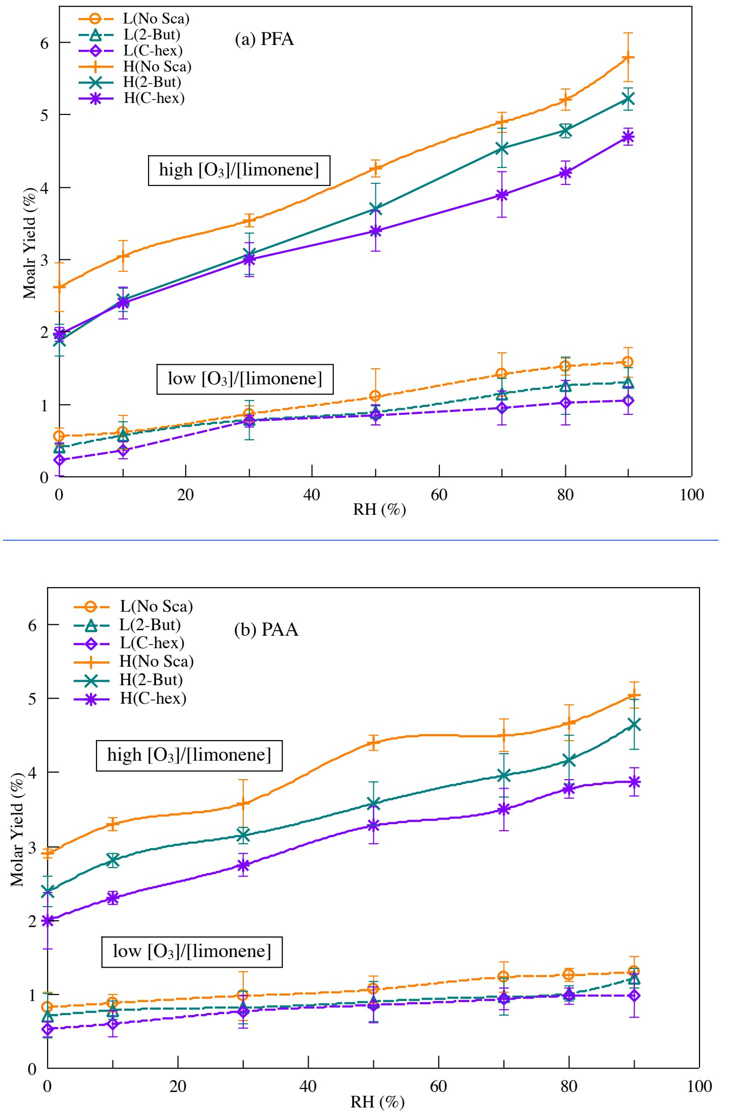
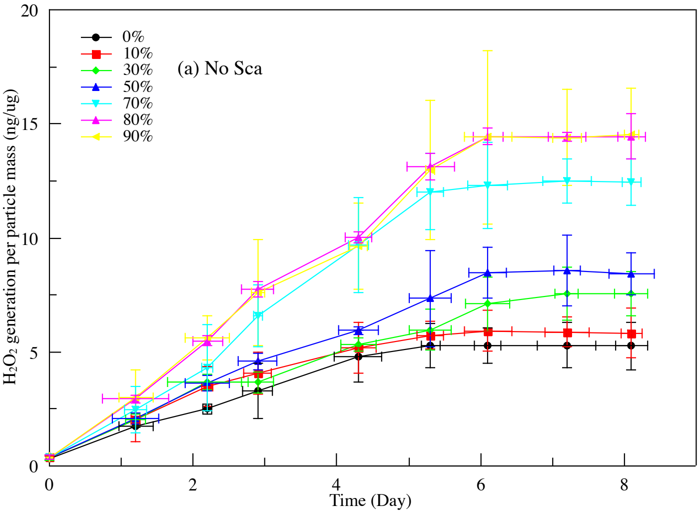
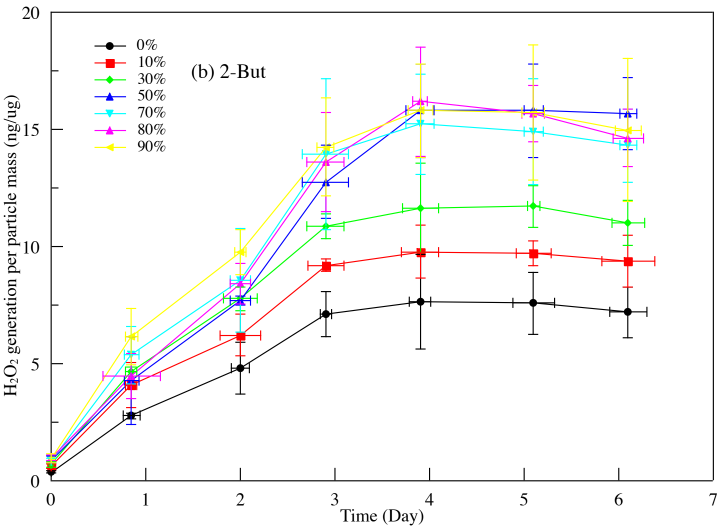


Figure 43. Dependence of (a) PFA yield and (b) PAA yield on relative humidity (RH) at low or high [O$_3$]/[limonene] ratio in the presence or absence of OH scavenger (2-butanol or cyclohexane). PFA, peroxyformic acid; PAA, peroxyacetic acid; L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.


(a) No Sca
0%
10%
30%
50%
70%
80%
90%
H$_2$O$_2$ generation per particle mass (ng/ug)
Time (Day)
0
5
10
15
20
0
2
4
6
8


(b) 2-But
0%
10%
30%
50%
70%
80%
90%
H$_2$O$_2$ generation per particle mass (ng/ug)
Time (Day)
0
5
10
15
20
0
1
2
3
4
5
6
7

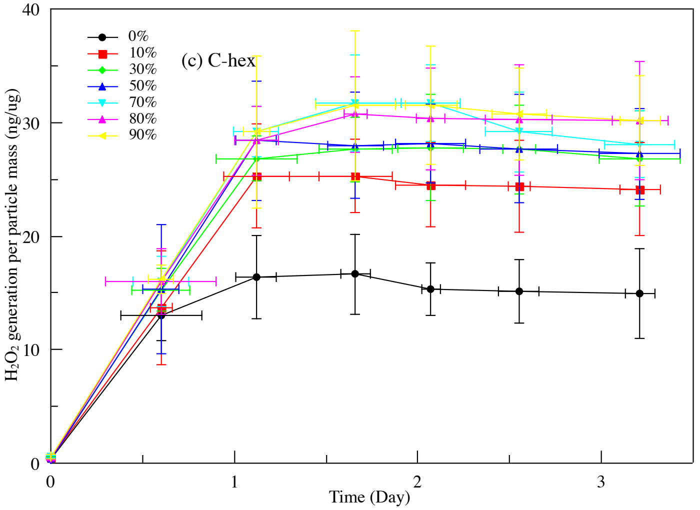
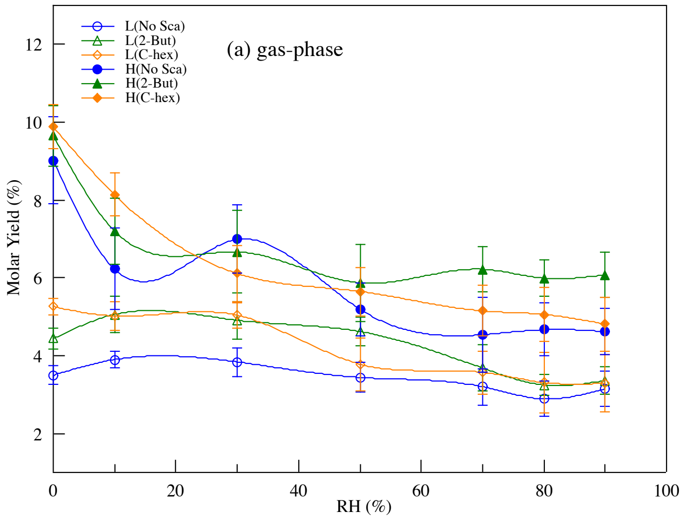
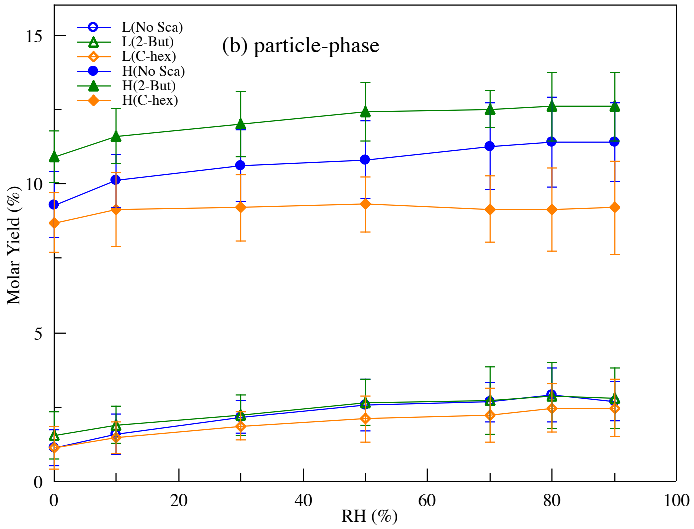


Figure ~~5~~4. Time profiles of H_2O_2 evolution per particle mass of different SOA formed (a) without OH scavenger, (b) with 2-butanol, and (c) with cyclohexane in the relative humidity (RH) range of 0–90% under high [O_3]/[limonene] ratio. H_2O_2, hydrogen peroxide; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.


L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
(a) gas-phase
Molar Yield (%)
2
4
6
8
10
12
0
20
40
60
80
100
RH (%)


L(No Sca)
L(2-But)
L(C-hex)
H(No Sca)
H(2-But)
H(C-hex)
(b) particle-phase
Molar Yield (%)
0
5
10
15
0
20
40
60
80
100
RH (%)

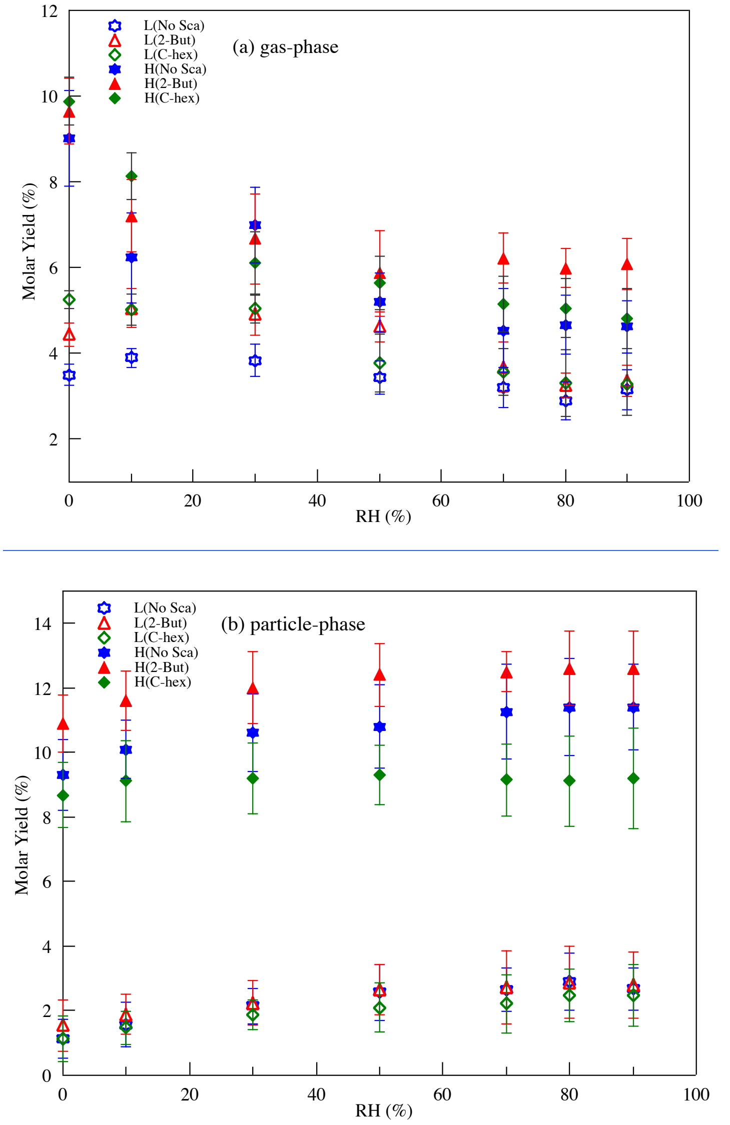


Figure ~~6~~5. The variation of (a) gas-phase and (b) particle-phase high-molecular-weight peroxides molar yields with relative humidity (RH) at low or high $[O_3]$/[limonene] ratio in the presence or absence of OH scavenger (2-butanol or cyclohexane). L, low ratio; H, high ratio; No-sca, none scavenger; 2-But, 2-butanol; C-hex, cyclohexane.

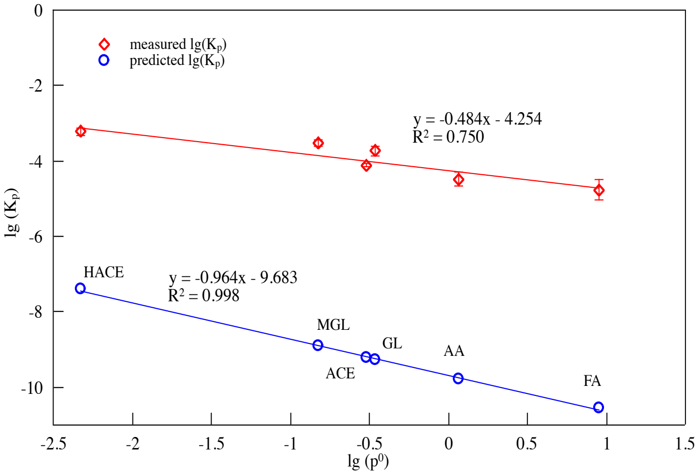


Figure ~~7~~6. The relationship of measured and predicted partitioning coefficients (K_p) versus vapor pressure (p^0) of carbonyls produced in limonene ozonolysis. HACE, hydroxyacetone; FA, formaldehyde; AA, acetaldehyde; ACE, acetone; GL, glyoxal; MGL, methylglyoxal.