

# ***Interactive comment on “Liquid-liquid phase separation in particles containing secondary organic material free of inorganic salts” by Mijung Song et al.***

**Mijung Song et al.**

[mijung.song@jbnu.ac.kr](mailto:mijung.song@jbnu.ac.kr)

Received and published: 8 August 2017

David Topping, Co-Editor of ACP

Dear David Topping,

Listed below are our responses to the comments from the reviewers of our manuscript. We thank the reviewers for carefully reading our manuscript and for their very helpful suggestions! For clarity and visual distinction, the referee comments or questions are listed here in black and are preceded by bracketed, italicized numbers (e.g. [1]). Authors' responses are in red below each referee statement with matching numbers (e.g.

[Printer-friendly version](#)

[Discussion paper](#)



[A1]).

Sincerely,

Allan Bertram Professor of Chemistry University of British Columbia

Response to Referee #1 (Reviewer comments in black text)

Summary: I have reviewed the paper “Liquid-liquid phase separation in particles containing secondary organic material free of inorganic salts” by the authors M. Song, P. Liu, S. T. Martin, and A. K. Bertram, which was submitted to *Atmos. Chem. Phys. Discuss.* This paper expands on previous work in which Renbaum-Wolff et al. found that in particles composed of secondary organic material (SOM) generated from a-pinene, liquid-liquid phase separation (LLPS) could occur at very high relative humidities (RH > 90%) in the absence of salt. Such a result may explain discontinuities in hygroscopicity observed below and above water saturation. This paper extends this research to SOM generated from b-caryophyllene, limonene, and toluene. LLPS is observed for the first two compounds, and not for the third. Overall, the data in the paper are of interest to atmospheric chemists and will make a good contribution to this journal. I have several major and minor comments, as indicated below, which should be addressed prior to publication.

Major comments are preceded by a “\*”.

[1] Abstract: \* O:C ranges overlap, which could be confusing to readers. Since only three systems are studied, it would be more accurate to state that you observed phase separation for BVOC oxidation products and not for toluene SOA. Plus, the O:C ratios obtained are averages over all SOA compounds formed in the experiment. LLPS may be observed due to a small concentration of compounds at very low O:C. To state a range may therefore be misleading.

[A1] Thank you for the comment. To avoid confusion, in the Abstract we will remove the sentence that mentions the O:C range. In addition, we will remove “average” from



## Interactive comment

the Abstract since LLPS can depend on both the average O:C and distribution of O:C.

Intro: [2] pg 2 line 8: Both low volatility and semivolatiles can partition to the particle phase

[A2] We will adjust this sentence to make it clear that semivolatiles can also partition to the particle phase.

[3] pg 2 line 24: LLPS can occur at O:C < 0.56 and does not occur at O:C > 0.8 according to your previous papers

[A3] The LLPS range will be corrected in the revised manuscript to make this point clear.

[4] pg 2 line 27: Peters is not cited. Perhaps you mean Petters et al.?

[A4] Correct! This will be revised in the final version.

[5] pg 2-3 lines 29,30,1: Change to “which could result in altered CCN properties” or similar, as the results from Renbaum-Wolff et al. are simulations of CCN data rather than experiments

[A5] We will revise as suggested.

[6] pg 3 lines 1-4: The results of Hodas et al. ACP 2016, 16, 12767-12792 should be discussed here.

[A6] We will include the reference and discussion in the revised manuscript.

[7] Materials and Methods: The error for the hygrometer (2.0% RH) seems high. Can the authors comment further about where this error is coming from (e.g. rate of RH ramp)?

[A7] The RH was determined from measurements of the temperature with a thermocouple and measurements of the dew point/frost point with a chilled mirror sensor (General Eastern, Canada). The RH was calibrated using the deliquescence RH for pure

[Printer-friendly version](#)

[Discussion paper](#)



ammonium sulfate particles (80% RH at 293 K, Martin, 2000). After calibration, the uncertainty of the hygrometer was  $\pm 2.0$  % RH based on the reproducibility of multiple deliquesce measurements. For clarification, this information will be added to the revised manuscript.

[8] Results and Discussion: \* pg 6: More evidence is needed to support the claim that LLPS is occurring by spinodal decomposition. Comparing to the data in Ciobanu et al. 2009, both the results in this paper and Renbaum-Wolff et al. 2016 clearly look like nucleation and growth (formation of small inclusions), as no schlieren are observed.

[A8] Spinodal decomposition is only clear in the movies provided in the Supplementary Material of this manuscript. Since this is not an important point in our paper, we will remove the two sentences from the paper that refer to spinodal decomposition.

[9] \* pg 6: It is reasonable to assume that the water-rich organic component is in the core of the particle and the more-insoluble organic is in the shell, but what evidence is there for this assignment?

[A9] We assume the core of the particle is water-rich, in part, because the size of the core decreases and eventually disappears as the RH decreases. The surface tension of water and the surface tensions of more-oxidized and less-oxidized organics are consistent with this assumption (Jasper, 1972). To address the referee's comment, we will add this information to the revised manuscript.

Jasper, J. J.: The surface tension of pure liquid compounds, J. Phys. And Chem. Ref. Data, vol 1, 841-1009, Doi: <http://dx.doi.org/10.1063/1.3253106>, 1972.

[10] \* pg 6: It is important to note that while core-shell structures are observed for these systems, without the substrate, they could be either partially engulfed or core-shell. It is less clear that there would be an effect of LLPS on CCN for these systems if they are partially engulfed in morphology.

[A10] We agree that a different morphology could result if the substrate is removed.

[Printer-friendly version](#)

[Discussion paper](#)



We will add this point to the manuscript to address the referee's comment.

[11] pg 7 line 4: The term “mixing” is more consistent with previous literature in this field (see for example Fig. 2 in the review from your group by You et al. in 2014). Later in this section, “MRH” is the mixing RH.

[A11] In the revised manuscript, we will replace “merging” with “mixing” as suggested.

[12] pg 7 line 8: Do you expect that particle mass concentrations will change O:C? Why is there a question that this will have an effect on LLPS?

[A12] Previous work has shown that the O:C of SOM can depend on the particle mass concentrations used to generate the SOM (E.g. Shilling et al. ACP, 2009, 771-782). To address this the referee's comments, this information will be added to the revised manuscript.

[13] pg 8 lines 15-23: SOA has a range of O:C values; the best that can be reported are average values. It could just be the very low O:C compounds that phase separate. It is therefore hard to understand how the systems used in this experiment can be compared to systems composed of only three components.

[A13] We completely agree that the spread of O:C values in the mixture is important, as well as the average O:C values. We will add the text below to the revised manuscript to make this clear.

“SOM have an average O:C and a spread (or distribution) of O:C values. Similar to SOM, systems containing two organics and water also have a spread in O:C and an average O:C. Hence, as a starting point to understanding LLPS in SOM, we considered previous studies that explored the miscibility gap in bulk solutions containing two organics and water (see Table 1 in Ganbaavale et al., 2015).”

[14] Implications: pg 9 lines 9-13: The Hodas et al. reference mentioned above should be added here.

Interactive comment

Printer-friendly version

Discussion paper



[A14] This reference will be included in the revised manuscript.

[15] Table 1&2: It would be nice to list average O:C for these exact systems, if known.

[A15] We did measure the O:C of the toluene-derived SOM studied in our manuscript. This information will be added to the Methods in the revised manuscript. The O:C of the other SOM was not measured by us; hence we rely on average O:C measurements from other studies using similar conditions. See [A16] below.

[16] \* Table 3: This table shows the minimum and maximum O:C range observed in the literature for SOM derived from these species, rather than O:C from the experiment run in this study. This is not ideal, but it may be due to experimental limitations. I worry, however, that O:C will change dramatically depending on how a system is oxidized and for how long. Are these comparable techniques to what you used? Do the numbers represent both chamber and flow tube studies? More information is needed here or in the text to understand how to interpret this data. Plus , table 3 & 4 are not consistent with one another (the highest O:C for alpha-pinene is 0.70 according to Table 4).

[A16] Regarding Table 3: The O:C of toluene-derived SOM investigated in this study was determined by us. This information will be added to the revised manuscript. The O:C of the other SOM investigated in our manuscript were not measured by us; hence, we rely on the O:C measurements from other studies. In the initial version of the manuscript, we included all previous O:C measurements. The referee is correct that O:C can change based on oxidation level and oxidation time. To address the referee's comments, in Table 3, we will only include O:C values from the literature that were determined using similar experimental conditions to those used in the LLPS studies. To illustrate the similarity in experimental conditions, in the revised manuscript (Table S1) we will list the experiment conditions used to generate the SOM in the LLPS studies and the experimental conditions used to generate the SOM in the O:C studies reported in Table 3.

Regarding Table 4: The average O:C values reported in Table 4 are based on mea-



surements in from the individual studies. The experimental conditions for the studies reported in Table 4 are not necessarily similar to the experimental conditions for the studies reported in Table 3, even if the same precursor volatile organic compound was used. For clarity, this information will be added to the revised manuscript.

[17] Figure 4: The first sentence of the legend is not clear. You are plotting SRH and MRH for the systems.

[A17] We will revise the sentence for clarify.

Please also note the supplement to this comment:

<https://www.atmos-chem-phys-discuss.net/acp-2017-408/acp-2017-408-AC1-supplement.pdf>

---

Interactive comment on *Atmos. Chem. Phys. Discuss.*, <https://doi.org/10.5194/acp-2017-408>, 2017.

[Printer-friendly version](#)

[Discussion paper](#)

