

Responses to Reviewer Comments:

We thank the reviewer for the constructive comments. The manuscript is revised based on the suggestions made and detailed responses to the reviewer are addressed as follows.

Reviewer #1

The revised version addresses the concerns of the reviewers, with two notable exceptions. First, the authors still do not correct for wall losses, but now clearly state this and mention that this will affect the reported yields. I'm willing to accept that, as long as they state the direction of this effect (see comment regarding line 272 below).

Author's comment

The wall loss might have an influence due to partitioning of the precursor compounds and products to the wall of the aerosol chamber resulting in lower SOA yields. To mention this in the manuscript the sentence (Line 270) was changed to "*Notably, the SOA yields of 1,3,5-TMB are not corrected for the wall loss to the surface of the aerosol chamber, which might possibly lead to an underestimation of the reported SOA yields.*".

Second, the authors make an argument that particle phase may influence the results, but cannot be predicted. This is generally true, but in my opinion some of the information provided in their responses to reviewers' comments would be very helpful to readers if included in the manuscript. Specifically, the differences between the inorganic components ammonium sulfate (solid at all RH used in the study) and ammonium bisulfate (liquid at all RH used in the study except <2% RH) are significant and should be mentioned. The authors could also mention the caveats (also given in their response to reviewers' comments) that the organic phase could lower the DRH of ammonium sulfate and cause it to deliquesce at 75% RH (but not likely at 0 or 50% RH), and that the formation of significant quantities of methylglyoxal oligomers could cause liquid aerosol particles (or liquid organic particle phases) to become semi-solid. Based on the work of Bertram et al.,⁽¹⁾ I think the authors can confidently state that at 0 and 50% RH, ammonium sulfate will be present in crystalline form along with an organic phase (of unknown viscosity), while at 50 and 75% RH crystalline ammonium bisulfate will not be present. If the authors included seed particle loadings, readers could interpret the ammonium sulfate data according to the organic / sulfate ratio framework of Bertram et al.⁽¹⁾

Author's comment

To clarify this some more information has been included in the manuscript.

Line 281: *“The RH value can have an influence on the phase state of the particles which, in turn, has an effect on the partitioning of the compounds into the particles as well as on particle-phase reactions (Ziemann, 2010; Saukko et al., 2012). The deliquescence RH of (NH₄)₂SO₄ and NH₄HSO₄ seed particles are known from literature to be 79% and 39%, respectively (Cziczo et al., 1997). Thus, pure NH₄HSO₄ seed particles are liquid at the applied RH values of 50% and 75% and solid at RH = 0% while (NH₄)₂SO₄ seed particles are solid over the whole RH range. Nevertheless, in the present study the particles are a mixture of inorganic and partitioned organic compounds. Organic compounds might change the phase state of the particles (Bertram et al., 2011) due to an influence on the deliquescence point (Andrews and Larson, 1993; Lightstone et al., 2000) as well as the hygroscopic behaviour of the particles (Lightstone et al., 2000; Prenni et al., 2003; Chen and Lee, 1999). Virtanen et al., 2010 postulated that particles are in an amorphous solid state if oligomeric compounds are present in the particles. Thus, it can be speculated that the particles in the present study containing a fraction of up to 8% of oligomeric compounds might be in an amorphous solid phase state. It was assumed, that the further reactions in the particle phase might be inhibited in solid particles (Saukko et al., 2012) thus, further oligomerisation or other reactions might become less effective after a certain fraction of oligomers exist in the particles. In addition, the partitioning of methylglyoxal monomers can be inhibited into solid particles (Saukko et al., 2012), which then might also lead to lower oligomer fractions in SOA. Thus, it might be possible that the phase state influences ΔM and the SOA yields.”*

Accordingly, the following references were included in the manuscript:

“Andrews, E., and Larson, S. M.: Effect of surfactant layers on the size changes of aerosol-particles as a function of relative-humidity, Environmental Science & Technology, 27, 857-865, 10.1021/es00042a007, 1993.”

“Bertram, A. K., Martin, S. T., Hanna, S. J., Smith, M. L., Bodsworth, A., Chen, Q., Kuwata, M., Liu, A.; You, Y., and Zorn, S. R.: Predicting the relative humidities of liquid-liquid phase separation, efflorescence, and deliquescence of mixed particles of ammonium sulfate, organic material, and water using the organic-to-sulfate mass ratio of the particle and the oxygen-to-carbon elemental ratio of the organic component, Atmospheric Chemistry and Physics, 11, 10995-11006, 10.5194/acp-11-10995-2011, 2011.”

“Chen, Y. Y., and Lee, W. M. G.: Hygroscopic properties of inorganic-salt aerosol with surface-active organic compounds, Chemosphere, 38, 2431-2448, 10.1016/s0045-6535(98)00436-6, 1999.”

“Cziczo, D. J., Nowak, J. B., Hu, J. H., and Abbatt, J. P. D.: Infrared spectroscopy of model tropospheric aerosols as a function of relative humidity: Observation of deliquescence and crystallization, Journal of Geophysical Research-Atmospheres, 102, 18843-18850, 10.1029/97jd01361, 1997.”

“Lightstone, J. M., Onasch, T. B., Imre, D., and Oatis, S.: Deliquescence, efflorescence, and water activity in ammonium nitrate and mixed ammonium nitrate/succinic acid microparticles, Journal of Physical Chemistry A, 104, 9337-9346, 10.1021/jp002137h, 2000.”

“Prenni, A. J., De Mott, P. J., and Kreidenweis, S. M.: Water uptake of internally mixed particles containing ammonium sulfate and dicarboxylic acids, Atmospheric Environment, 37, 4243-4251, 10.1016/s1352-2310(03)00559-4, 2003.”

“Virtanen, A., Joutsensaari, J., Koop, T., Kannosto, J., Yli-Pirila, P., Leskinen, J., Makela, J. M., Holopainen, J. K., Poschl, U., Kulmala, M., Worsnop, D. R., and Laaksonen, A.: An amorphous solid state of biogenic secondary organic aerosol particles, Nature, 467, 824-827, 10.1038/nature09455, 2010.”

Other comments:

Line 75: The statement “The method is applicable for all oligomeric compounds, which can be decomposed into methylglyoxal monomers” is misleading as currently stated. It would be clearer to state “The method is applicable for any oligomeric compound which can be decomposed into methylglyoxal monomers.”

Author’s comment

The author’s agree with the reviewer and changed the sentence to “*The method is applicable for any oligomeric compounds, which can be decomposed into methylglyoxal monomers during the heating process at a temperature of 100 °C.*”

Line 133: Technically, the filter extract that was not heated is not measuring only methylglyoxal monomers, but also any compound that can be hydrolyzed at room temperature into methylglyoxal monomers. This likely includes some acetal oligomers. Also, the authors should make clear at some point in the manuscript that “methylglyoxal monomers” includes singly and doubly hydrated methylglyoxal.

Author’s comment

The author's agree with the reviewer. The sentence was changed to *“The second half of the filters was used to quantify monomeric methylglyoxal including also singly and doubly hydrated methylglyoxal, which can be hydrolysed at room temperature into monomers.”*

Line 394: There also might be oligomers originating from methylglyoxal monomers that are not heat-decomposable back to methylglyoxal monomers, which would also not be determined in the present study.

To clarify this, the sentence was changed to *“[...] thus there might be not heat-decomposable methylglyoxal oligomers or oligomers originating from other monomers [...]”*

Technical corrections:

Line 20: The phrase “emitted hydrocarbons contributing with up to ...“ would be clearer if written as “emitted hydrocarbons, contributing up to ...”

The sentence was changed to *“[...] fraction of the emitted hydrocarbons, contributing up to 52% [...]”*

Line 24: “low-volatile” should be “low-volatility”

The sentence was changed to *“[...] leads to low-volatility oxidation products [...]”*

Line 127 and 271: I think that “Noticeably” should be “Notably”

“Noticeably” was changed to *“Notably”* in Line 127 and 271.

Line 148: it would be helpful to give essential parameters of this method here.

“The filters were prepared according to the method described by Rodigast et al. (2015).” was changed to *“The filters were prepared according to the method described by Rodigast et al. (2015), thus 300 μ L of an aqueous PFBHA solution (5 mg mL⁻¹) was added to the filter extracts. After a derivatisation time of 24 hours at room temperature the derivatised methylglyoxal monomers were extracted.”*

Line 272: The phrase “...might have an influence on the reported SOA yields.” would be more informative as “...might exert a downward influence on the reported SOA yields.”

The sentence (Line 270) was changed to *“Notably, the SOA yields of 1,3,5-TMB are not corrected for the wall loss to the surface of the aerosol chamber, which might possibly lead to an underestimation of the reported SOA yields.”*

Line 299: The phrase “... controversially discussed in the literature” could be changed to a single word “... controversial”

The sentence was changed to *“[...] on SOA formation is controversial [...]”*.

Table 4: It would be good to include the results of the present study in Table 4.

The results of the present study were included in Table 4.

The following changes were made to the manuscript

Line 20: *“Aromatic compounds represent a large fraction of the emitted hydrocarbons contributing with up to 52% to the total non-methane hydrocarbon mass at an industrial dominated site in China (Liu et al., 2008).”* was changed to *“Aromatic compounds represent a large fraction of the emitted hydrocarbons, contributing up to 52% to the total non-methane hydrocarbon mass at an industrial dominated site in China (Liu et al., 2008).”*

Line 23: *“The gas-phase oxidation of 1,3,5-TMB leads to low-volatile oxidation products, which partition into the particle phase and form secondary organic aerosol (SOA).”* was changed to *“The gas-phase oxidation of 1,3,5-TMB leads to low-volatility oxidation products, which partition into the particle phase and form secondary organic aerosol (SOA).”*

Line 75: *“The method is applicable for all oligomeric compounds, which can be decomposed into methylglyoxal monomers during the heating process at a temperature of 100 °C.”* was changed to *“The method is applicable for any oligomeric compounds, which can be decomposed into methylglyoxal monomers during the heating process at a temperature of 100 °C.”*

Line 127: *“Noticeably, the extraction efficiency was not investigated in the present study, thus it is not known.”* was changed to *“Notably, the extraction efficiency was not investigated in the present study, thus it is not known.”*

Line 146: *“The second half of the filters was used to quantify monomeric methylglyoxal.”* was changed to *“The second half of the filters was used to quantify monomeric methylglyoxal”*

including also singly and doubly hydrated methylglyoxal, which can be hydrolysed at room temperature into monomers.”

Line 146: *“The filters were prepared according to the method described by Rodigast et al. (2015).”* was changed to *“The filters were prepared according to the method described by Rodigast et al. (2015), thus 300 μL of an aqueous PFBHA solution (5 mg mL^{-1}) was added to the filter extracts. After a derivatisation time of 24 hours at room temperature the derivatised methylglyoxal monomers were extracted.”*

Line 270: *“Noticeably, the SOA yields of 1,3,5-TMB are not corrected for the wall loss to the surface of the aerosol chamber, which might have an influence on the reported SOA yields.”* was changed to *“Notably, the SOA yields of 1,3,5-TMB are not corrected for the wall loss to the surface of the aerosol chamber, which might possibly lead to an underestimation of the reported SOA yields.”*

Line 281: *“ The RH value can have an influence on the phase state of the particles while the phase state has an effect on the partitioning of the compounds into the particles and the particle-phase reactions (Ziemann, 2010; Saukko et al., 2012). In the present study the particles are a mixture of inorganic and partitioned organic compounds, thus the phase state is not known but it might be possible that the phase state influence ΔM and the SOA yields.”* was changed to *“The RH value can have an influence on the phase state of the particles which, in turn, has an effect on the partitioning of the compounds into the particles as well as on particle-phase reactions (Ziemann, 2010; Saukko et al., 2012). The deliquescence RH of $(\text{NH}_4)_2\text{SO}_4$ and NH_4HSO_4 seed particles are known from literature to be 79% and 39%, respectively (Cziczo et al., 1997). Thus, pure NH_4HSO_4 seed particles are liquid at the applied RH values of 50% and 75% and solid at $\text{RH} = 0\%$ while $(\text{NH}_4)_2\text{SO}_4$ seed particles are solid over the whole RH range. Nevertheless, in the present study the particles are a mixture of inorganic and partitioned organic compounds. Organic compounds might change the phase state of the particles (Bertram et al., 2011) due to an influence on the deliquescence point (Andrews and Larson, 1993; Lightstone et al., 2000) as well as the hygroscopic behaviour of the particles (Lightstone et al., 2000; Prenni et al., 2003; Chen and Lee, 1999). Virtanen et al., 2010 postulated that particles are in an amorphous solid state if oligomeric compounds are present in the particles. Thus, it can be speculated that the particles in the present study containing a fraction of up to 8% of oligomeric compounds might be in an amorphous solid phase state. It was assumed, that the further reactions in the particle phase might be inhibited*

in solid particles (Saukko et al., 2012) thus, further oligomerisation or other reactions might become less effective after a certain fraction of oligomers exist in the particles. In addition, the partitioning of methylglyoxal monomers can be inhibited into solid particles (Saukko et al., 2012), which then might also lead to lower oligomer fractions in SOA. Thus, it might be possible that the phase state influences ΔM and the SOA yields.”

Line 298: “*The influence of RH on SOA formation is controversially discussed in the literature (Hennigan et al., 2008; Fick et al., 2003; Edney et al., 2000; Saxena and Hildemann, 1996; Baker et al., 2001; Hasson et al., 2001, Cocker et al., 2001)*” was changed to “*The influence of RH on SOA formation is controversial (Hennigan et al., 2008; Fick et al., 2003; Edney et al., 2000; Saxena and Hildemann, 1996; Baker et al., 2001; Hasson et al., 2001, Cocker et al., 2001).*”

Line 392: “*In the present study only heat-decomposable methylglyoxal oligomers were quantified, thus there might be oligomers originating from other monomers than methylglyoxal, which were not determined in the present study leading to lower oligomer fractions of ΔM compared to the literature studies.*” was changed to “*In the present study only heat-decomposable methylglyoxal oligomers were quantified, thus there might be not heat-decomposable methylglyoxal oligomers or oligomers originating from other monomers than methylglyoxal, which were not determined in the present study leading to lower oligomer fractions of ΔM compared to the literature studies.*”

The reference (Line 425) “*Andrews, E., and Larson, S. M.: Effect of surfactant layers on the size changes of aerosol-particles as a function of relative-humidity, Environmental Science & Technology, 27, 857-865, 10.1021/es00042a007, 1993.*” was included in the manuscript.

The reference (Line 447) “*Bertram, A. K., Martin, S. T., Hanna, S. J., Smith, M. L., Bodsworth, A., Chen, Q., Kuwata, M., Liu, A.; You, Y., and Zorn, S. R.: Predicting the relative humidities of liquid-liquid phase separation, efflorescence, and deliquescence of mixed particles of ammonium sulfate, organic material, and water using the organic-to-sulfate mass ratio of the particle and the oxygen-to-carbon elemental ratio of the organic component, Atmospheric Chemistry and Physics, 11, 10995-11006, 10.5194/acp-11-10995-2011, 2011.*” was included in the manuscript.

The reference (Line 460) “Chen, Y. Y., and Lee, W. M. G.: *Hygroscopic properties of inorganic-salt aerosol with surface-active organic compounds*, *Chemosphere*, 38, 2431-2448, 10.1016/s0045-6535(98)00436-6, 1999.” was included in the manuscript.

The reference (Line 466) “Cziczo, D. J., Nowak, J. B., Hu, J. H., and Abbatt, J. P. D.: *Infrared spectroscopy of model tropospheric aerosols as a function of relative humidity: Observation of deliquescence and crystallization*, *Journal of Geophysical Research-Atmospheres*, 102, 18843-18850, 10.1029/97jd01361, 1997.” was included in the manuscript.

The reference (Line 574) “Lightstone, J. M., Onasch, T. B., Imre, D., and Oatis, S.: *Deliquescence, efflorescence, and water activity in ammonium nitrate and mixed ammonium nitrate/succinic acid microparticles*, *Journal of Physical Chemistry A*, 104, 9337-9346, 10.1021/jp002137h, 2000.” was included in the manuscript.

The reference (Line 631) “Prenni, A. J., De Mott, P. J., and Kreidenweis, S. M.: *Water uptake of internally mixed particles containing ammonium sulfate and dicarboxylic acids*, *Atmospheric Environment*, 37, 4243-4251, 10.1016/s1352-2310(03)00559-4, 2003.” was included in the manuscript.

The reference (Line 714) “Virtanen, A., Joutsensaari, J., Koop, T., Kannosto, J., Yli-Pirila, P., Leskinen, J., Makela, J. M., Holopainen, J. K., Poschl, U., Kulmala, M., Worsnop, D. R., and Laaksonen, A.: *An amorphous solid state of biogenic secondary organic aerosol particles*, *Nature*, 467, 824-827, 10.1038/nature09455, 2010.” was included in the manuscript.

A quantification method for heat-decomposable methylglyoxal oligomers and its application on 1,3,5-trimethylbenzene SOA

Maria Rodigast, Anke Mutzel, and Hartmut Herrmann*

Leibniz Institute for Tropospheric Research (TROPOS), Atmospheric Chemistry Dept. (ACD), Permoserstr. 15, D-04318 Leipzig, Germany

*Corresponding author. Tel: +49-341-2717-7024; fax: +49-341-271799-7024.

E-mail address: herrmann@tropos.de

For submission to: Atmospheric Chemistry and Physics

First submitted on July 02nd 2016, revised August 09th, 2016

1 Abstract

2 Methylglyoxal forms oligomeric compounds in the atmospheric aqueous particle phase, which could establish
3 a significant contribution to the formation of aqueous secondary organic aerosol (aqSOA). Thus-far, no
4 suitable method for the quantification of methylglyoxal oligomers is available despite the great effort spent for
5 structure elucidation. In the present study a simplified method was developed to quantify heat-decomposable
6 methylglyoxal oligomers as a sum parameter. The method is based on the thermal decomposition of oligomers
7 into methylglyoxal monomers. Formed methylglyoxal monomers were detected using PFBHA (o-(2,3,4,5,6-
8 pentafluorobenzyl)hydroxylamine hydrochloride) derivatisation and gas chromatography/mass spectrometry
9 (GC/MS) analysis. The method development was focused on the heating time (varied between 15 and 48
10 hours), pH during the heating process (pH = 1 - 7), and heating temperature (50°C, 100°C). The optimised
11 values of these method parameters are presented.

12 The developed method was applied to quantify heat-decomposable methylglyoxal oligomers formed during
13 the OH-radical oxidation of 1,3,5-trimethylbenzene (TMB) in the Leipziger aerosol chamber (LEAK).
14 Oligomer formation was investigated as a function of seed particle acidity and relative humidity. A fraction of
15 heat-decomposable methylglyoxal oligomers of up to 8% in the produced organic particle mass was found,
16 highlighting the importance of those oligomers formed solely by methylglyoxal for SOA formation. Overall,
17 the present study provides a new and suitable method for quantification of heat-decomposable methylglyoxal
18 oligomers in the aqueous particle phase.

19 **1. Introduction**

20 Aromatic compounds represent a large fraction of the emitted hydrocarbons, contributing up to 52% to the
21 total non-methane hydrocarbon mass at an industrial dominated site in China (Liu et al., 2008). One of these
22 aromatic compounds is 1,3,5-trimethylbenzene (TMB), which was measured in the gas phase in concentrations
23 ranging from 0.7 to 40.6 $\mu\text{g m}^{-3}$ (Gee and Sollars, 1998; Khoder, 2007). The gas-phase oxidation of 1,3,5-
24 TMB leads to low-volatility oxidation products, which partition into the particle phase and form secondary
25 organic aerosol (SOA). Oxidation products of 1,3,5-TMB were investigated in a number of literature studies
26 (e. g. Huang et al., 2015; Baltensperger et al., 2005; Kalberer et al., 2004; Kalberer et al., 2006; Paulsen et al.,
27 2005; Healy et al., 2008; Cocker et al., 2001; Smith et al., 1999; Metzger et al., 2008; Wyche et al., 2009; Yu
28 et al., 1997). Methylglyoxal was found as one of the most important oxidation product (Metzger et al., 2008;
29 Healy et al., 2008; Cocker et al., 2001; Smith et al., 1999; Wyche et al., 2009; Baltensperger et al., 2005;
30 Rickard et al., 2010; Kalberer et al., 2004; Yu et al., 1997; Kleindienst et al., 1999; Müller et al., 2012; Nishino
31 et al., 2010; Hamilton et al., 2003; Tuazon et al., 1986; Bandow and Washida, 1985; Lim and Turpin, 2015)
32 contributing with a fraction of up to 2% to the particle mass (Healy et al., 2008; Cocker et al., 2001).
33 Methylglyoxal has often been described to form oligomeric compounds in the aqueous particle phase (see, e.
34 g. Herrmann et al., 2015 for an overview; De Haan et al., 2009; Kalberer et al., 2004; Loeffler et al., 2006;
35 Zhao et al., 2006; Sareen et al., 2010; Altieri et al., 2008), which are supposed to play an important role in the
36 formation of aqueous secondary organic aerosols (aqSOA; e. g. Kalberer et al., 2004).

37 In general, oligomeric compounds can be formed in the aqueous particle phase through aldol
38 condensation (e. g. Tilgner and Herrmann, 2010; Sareen et al., 2010; Sedehi et al., 2013; Krizner et al., 2009;
39 Barsanti and Pankow, 2005; De Haan et al., 2009; Yasmeen et al., 2010), acetal/hemiacetal formation (Kalberer
40 et al., 2004 and Yasmeen et al., 2010), esterification (Altieri et al., 2008; Sato et al., 2012; Tan et al., 2010; De
41 Haan et al., 2011; Sedehi et al., 2013), imine formation (Altieri et al., 2008; Sato et al., 2012; Tan et al., 2010;
42 De Haan et al., 2011; Sedehi et al., 2013), hydrolysis of epoxides (Paulot et al., 2009; Surratt et al., 2010),
43 polymerisation, and radical – radical reactions (Schaefer et al., 2015; Tan et al., 2012; Lim et al., 2013).
44 During the last decade, huge efforts were undertaken to detect and identify oligomeric compounds. As it can
45 be seen in Table 1, a number of mass spectrometric methods were used including (matrix assisted) laser
46 desorption/ionisation mass spectrometry (MALDI-MS, LDI-MS), Fourier transform ion cyclotron resonance
47 mass spectrometry (FT-ICR-MS), electrospray ionisation mass spectrometry and electrospray ionisation
48 tandem mass spectrometry (ESI/MS, ESI/MS/MS). In addition, spectroscopic methods like UV/Vis
49 (ultraviolet-visible spectroscopy), FTIR (Fourier transform infrared spectroscopy), and NMR (nuclear
50 magnetic resonance spectroscopy) analysis were used for identification.

51 Despite the past effort for structure elucidation of oligomeric compounds a suitable quantification
52 method is not available. Mostly, an overall contribution of oligomers to the particle mass was determined
53 using, e.g., a volatility tandem differential mobility analyser (VTDMA). Kalberer et al. (2004) determined an
54 oligomer contribution of 50% to the particle mass formed by the photooxidation of 1,3,5-TMB. In a further
55 experiment, oligomer mass fractions of 80% and 90% were determined with a VTDMA-based approach for
56 1,3,5-TMB and α -pinene (Kalberer et al., 2006). Alfarrá et al. (2006) investigated the photooxidation of 1,3,5-
57 TMB and found an increase of the oligomer fraction of 3.1 and 3.7% hour^{-1} . A particulate oligomer fraction of
58 50% was reported for 1,3,5-TMB and α -pinene by Baltensperger et al. (2005). Dommen et al. (2006) detected

59 a contribution of oligomers to the organic particle mass increasing from 27% to 44% in the first 5 hours of the
60 photooxidation of isoprene. Nguyen et al. (2011) investigated oligomers from isoprene photooxidation with
61 ESI-MS and nano-DESI-MS (nanospray desorption electrospray ionisation) connected to a high resolution
62 linear ion trap (LTQ-) orbitrap. They calculated an oligomer fraction of 80 – 90%.

63 De Haan et al. (2009) estimated the oligomer fraction formed by methylglyoxal in the aqueous phase
64 with NMR concluding 37% of methylglyoxal are dimers and oligomers. In another approach it was estimated
65 that after 4 days in aqueous particles containing amino acids, 15% of the carbonyl compounds are oligomers
66 (Noziere et al., 2007). Contrary, with ammonium sulfate particles 30% of the carbonyl compounds are
67 converted into oligomers (Noziere et al., 2007). These estimations are based on the rate constants for oligomer
68 formation which were determined in the study by Noziere et al. (2007). Besides these estimations,
69 quantification of oligomeric compounds was also conducted using surrogate compounds (Surratt et al., 2006,
70 Zappoli et al., 1999, and Gao et al., 2004) or synthesised authentic standards (Birdsall et al., 2013).

71 In summary, a variety of methods exists which quantify the fraction of oligomeric compounds derived
72 from methylglyoxal, but the results are contradicting due to the lack of a suitable method for quantification
73 and second, due to different reaction conditions used in the studies. Thus, the present study presents a
74 fundamental approach for a reliable quantification of methylglyoxal oligomers in laboratory-generated SOA.
75 **The method is applicable for any oligomeric compounds, which can be decomposed into methylglyoxal**
76 **monomers during the heating process at a temperature of 100 °C.** As the oligomerisation mechanisms leading
77 to the quantified oligomeric compounds are not known, it cannot be specified if the oligomers are reversibly
78 or irreversibly formed. In addition, it cannot be excluded that there exist oligomers, which are not
79 decomposable into their methylglyoxal monomers through the heating process. Thus, the quantified oligomers
80 are termed as heat-decomposable methylglyoxal oligomers.

81

82 **2. Experimental**

83 2.1 Chemicals and standards

84 1,3,5-TMB ($\geq 99.8\%$), hydrochloric acid (37%), tetramethylethylene (99%), and sodium hydroxide
85 (50 – 52%) were obtained from Sigma-Aldrich (Hamburg, Germany). *O*-(2,3,4,5,6-pentafluorobenzyl)-
86 hydroxylamine hydrochloride ($\geq 99\%$), methylglyoxal (40% in water), and ammonium hydrogensulfate (98%)
87 were purchased from Fluka (Hamburg, Germany). Sulfuric acid (98%) was obtained from Merck KGaA
88 (Darmstadt, Germany). Dichloromethane (Chromasolv 99.8%) was obtained from Riedel-de Haen (Seelze,
89 Germany) and ammonium sulfate (99.5%) was purchased from Carl Roth (Karlsruhe, Germany). Ultrapure
90 water was used to prepare the seed particle solutions, the authentic standards, and to extract the filter samples
91 (Milli-Q gradient A 10, 18.2 M Ω cm, 3 ppb TOC, Millipore, USA).

92

93 2.2 Chamber experiments

94 The OH-radical oxidation of 1,3,5-TMB was investigated in the LEipzigiger AerosolKammer (LEAK). A
95 detailed description of the aerosol chamber can be found elsewhere (Mutzel et al., 2016). The conditions of
96 the experimental runs are summarised in Table 2. The experiments were conducted in the presence of
97 ammonium bisulfate particles or ammonium sulfate particles mixed with sulfuric acid to achieve different seed
98 acidities. In order to investigate OH-radical oxidation of 1,3,5-TMB at low NO_x levels (< 1 ppb) and under

99 dark conditions the ozonolysis of tetramethylethylene (TME) was used as OH-radical source (Berndt and
100 Böge, 2006). The cycloaddition of ozone to TME yields a primary ozonide, which reacts further and forms a
101 stabilised Criegee Intermediate (sCI). The sCI decomposes via the hydroperoxide channel, leading to the
102 formation of OH radicals (Gutbrod et al., 1996) with a yield of 0.92 ± 0.08 (Berndt and Böge, 2006). O₃ was
103 produced by UV irradiation of O₂ with an O₂ flow rate of 5 L minute⁻¹. It was injected at the beginning of the
104 experiments and ≈ 26 ppbv of TME was introduced into the aerosol chamber in steps of 15 minutes. 1,3,5-
105 TMB (≈ 92 ppb) was injected into the aerosol chamber using a microliter-syringe. The oxidation of 1,3,5-TMB
106 was studied at relative humidities (RH) between $\approx 0\%$ and 75% adjusted by flushing the aerosol chamber with
107 humid or dry air. The consumption of the precursor compound (Δ HC) was monitored over a reaction time of
108 90 minutes with a proton-transfer-reaction time-of-flight mass spectrometer (PTR-TOF MS; 8000, IONICON
109 Analytik, Innsbruck, Germany). The volume size distribution of the seed particles was measured with a
110 scanning mobility particle sizer (SMPS; 3010; TSI, USA). An average density of 1 g cm⁻³ was used to calculate
111 the increase of the organic particle mass (Δ M). To collect the particle phase after the experiments, 1.2 m³ of
112 the chamber volume was sampled on a PTFE filter (borosilicate glass fiber filter coated with fluorocarbon, 47
113 mm in diameter, PALLFLEX T60A20, PALL, NY, USA) connected to a XAD-4 coated denuder (URG-2000-
114 30B5, URG Corporation, Chapel Hill, NC, USA) (Kahnt et al., 2011) to avoid artefacts caused by adsorption
115 of gas-phase organic compounds onto the filter.

116

117 2.3 Sample preparation

118 For method development, PTFE filter samples from aerosol chamber experiments were used. The following
119 method parameters were investigated: heating time, pH during the heating process, and heating temperature
120 (Table 3). Filter samples from the same experiments were used for the optimisation of the respective method
121 parameters.

122

123 2.3.1 Filter extraction

124 Two halves of the PTFE filters were cut into small pieces. Each filter half was extracted separately with 1 mL
125 H₂O for 30 minutes using an orbital shaker (700 rpm, revolutions per minutes). They were shaken again
126 separately with 1 mL H₂O for 30 minutes and flushed at the end of the procedure with 1 mL H₂O resulting in
127 two 3 mL extracts. **Notably, the extraction efficiency was not investigated in the present study, thus it is not**
128 **known.** Water was used as extracting reagent, because organic solvents like methanol and acetonitrile have
129 lower boiling points than water, thus lower heating temperatures can be applied for the decomposition of the
130 oligomers. Besides this, organic solvents like methanol and acetonitrile are miscible with dichloromethane,
131 thus an extraction of the derivatised methylglyoxal with dichloromethane prior the injection into GC/MS would
132 not be possible. The extract of one half of the filter was used for oligomer measurements (extract 1) and with
133 the second one methylglyoxal monomers were quantified which were not a building block of oligomers
134 (extract 2). A detailed description of the derivatisation procedure can be found in Rodigast et al. (2015).

135

136 2.3.2 Derivatisation procedure

137 *Extract 1 – heat-decomposable methylglyoxal oligomers*

138 For quantification of heat-decomposable methylglyoxal oligomers the extract was acidified and heated to
139 decompose the oligomeric bonds. The pH was adjusted with hydrochloric acid (37%) or sodium hydroxide
140 (1 mol L^{-1}) to pH = 1, 3, 5, and 7 while heating temperatures of 50°C and 100°C were investigated. For the
141 derivatisation of the formed monomeric methylglyoxal, $300 \mu\text{L}$ of *o*-(2,3,4,5,6-
142 pentafluorobenzyl)hydroxylamine hydrochloride (PFBHA, 5 mg mL^{-1}) was added to the sample solution after
143 2 minutes of the heating process. Different heating times were tested varying between 15 hours and 48 hours.
144 After the derivatisation was complete, the extracts were allowed to cool down to room temperature.

145

146 *Extract 2 – Methylglyoxal monomer*

147 The second half of the filters was used to quantify monomeric methylglyoxal including also singly and doubly
148 hydrated methylglyoxal, which can be hydrolysed at room temperature into monomers. The filters were
149 prepared according to the method described by Rodigast et al. (2015), thus $300 \mu\text{L}$ of an aqueous PFBHA
150 solution (5 mg mL^{-1}) was added to the filter extracts. After a derivatisation time of 24 hours at room
151 temperature the derivatised methylglyoxal monomers were extracted.

152

153 2.3.3 Extraction for GC/MS analysis

154 After derivatisation of both filter extracts (extract 1 and 2), derivatised methylglyoxal was extracted at pH = 1
155 for 30 minutes with $250 \mu\text{L}$ of dichloromethane using an orbital shaker (1500 rpm; Rodigast et al., 2015). $1 \mu\text{L}$
156 of the organic phase was injected into GC/MS for analysis. The measurements were repeated for three times
157 to ensure reliable GC/MS signals.

158 For quantification, a 5-point calibration was performed at the beginning of each chromatographic run using a
159 standard solution of methylglyoxal in a concentration range of 0.13 to $8 \mu\text{mol L}^{-1}$.

160

161 2.4 Instrumentation

162 The samples were analysed using a GC System (6890 Series Agilent Technologies, Frankfurt, Germany)
163 coupled with an electron ionisation quadrupole mass spectrometer in splitless mode with an inlet temperature
164 of 250°C (Agilent 5973 Network mass selective detector, Frankfurt, Germany). The derivatives were separated
165 with a HP-5MS UI column (Agilent J & W GC columns, $30 \text{ m} \times 0.25 \text{ mm} \times 0.25 \mu\text{m}$) using the following
166 temperature program: 50°C isothermal for 2 minutes and elevated to 230°C ($10^\circ\text{C minute}^{-1}$). The temperature
167 of 230°C was held constant for 1 minute and ended with 320°C for 10 minutes, thus the method has a run time
168 of 36 minutes.

169

170 3. Results

171 Hastings et al. (2005) investigated the influence of the temperature of the GC inlet on the detection of
172 oligomeric compounds. These authors concluded that oligomers decompose into monomer building blocks at
173 higher inlet temperatures ($\geq 120^\circ\text{C}$) which caused problems for oligomer quantification. In the present study a
174 quantification method is proposed to decompose heat-decomposable oligomers into methylglyoxal monomers
175 due to heating, acidification and PFBHA derivatisation prior GC/MS injection.

176 (E) and (Z) isomers of methylglyoxal were formed during PFBHA derivatisation resulting in two peaks in the
177 GC/MS chromatogram. For quantification, the sum of these peaks was used to avoid an over- or
178 underestimation of methylglyoxal due to variations of the isomer peak ratio during the heating process.

179

180 **3.1 Method development**

181 *Influence of heating time*

182 The influence of the heating time was examined with PTFE filters which were sampled after OH-radical
183 oxidation of 1,3,5-TMB at RH = 75% in the presence of NH₄HSO₄ seed particles (experiment #3). To
184 investigate the effect of the heating time on the decomposition of the heat-decomposable oligomeric
185 compounds, the aqueous filter extracts (extract 1) were acidified to pH = 1 and heated to 100°C for
186 15 – 48 hours. The results were compared to the unheated aqueous filter extracts (extract 2) to determine the
187 increase of methylglyoxal concentration due to decomposition of the heat-decomposable oligomers.
188 Additionally, a 6.25 μmol L⁻¹ standard solution of methylglyoxal was acidified and heated for different times
189 to exclude an effect of the heating process on the derivatisation. The results are illustrated in Fig. 1a.

190 The highest methylglyoxal concentration can be found after a heating time of 24 hours. The methylglyoxal
191 concentration was about six times higher ($c = 1.82 \pm 0.14 \mu\text{mol L}^{-1}$) in comparison to the unheated filter extract
192 ($c = 0.29 \pm 0.01 \mu\text{mol L}^{-1}$). To exclude that the higher methylglyoxal concentrations were only a result of a
193 better PFBHA derivatisation during heating, a methylglyoxal standard solution was also heated for 24 hours
194 (Fig. 1b). A methylglyoxal concentration of $c = 5.32 \pm 0.05 \mu\text{mol L}^{-1}$ was found, which corresponds to a
195 recovery of $\approx 85\%$. Thus, an effect of the heating process on the derivatisation can be excluded indicating that
196 the higher methylglyoxal concentration was caused by decomposition of heat-decomposable oligomers into
197 monomers.

198 One filter extract was heated for 15 hours and allowed to stand at room temperature for 9 hours
199 (sample a) to reach a total derivatisation time of 24 hours (as it was optimised for PFBHA derivatisation by
200 Rodigast et al., 2015). To exclude reoligomerisation processes of methylglyoxal in sample a, one filter extract
201 was heated for 15 hours and measured directly after the heating process (sample b). As it can be seen in Fig. 1a
202 both filter samples (sample a and b) showed lower methylglyoxal concentrations than after heating for 24
203 hours. The lower methylglyoxal concentration of sample b might be caused by an incomplete derivatisation
204 due to the immediate measurement of the filter extract after 15 hours heating time. In comparison, the lower
205 concentration in sample a might be caused by reoligomerisation of methylglyoxal.

206 To probe this hypothesis a 6.25 μmol L⁻¹ methylglyoxal standard was heated for 15 hours and
207 measured immediately (like sample a) or, alternatively, was allowed to stand at room temperature for 9 hours
208 (like sample b). In Fig. 1b a lower methylglyoxal concentration can be observed for the immediately measured
209 sample ($3.11 \pm 0.20 \mu\text{mol L}^{-1}$) compared to the sample after 9 hours at room temperature
210 ($5.84 \pm 0.27 \mu\text{mol L}^{-1}$). Thus, it can be concluded that a derivatisation time of 24 hours is needed for a complete
211 derivatisation despite the heating process. This supports the hypothesis that methylglyoxal monomers were not
212 completely derivatised, if the filter sample was heated for 15 hours and directly measured (sample b). Based
213 on the incomplete derivatisation after 15 hours heating time, methylglyoxal monomers are able to react again
214 under oligomer formation during the 9 hours at room temperature. The rate constants are reported to be
215 $k = 5 \times 10^{-6} \text{ M}^{-1} \text{ minutes}^{-1}$ for ammonium ion catalysed and $k \leq 1 \times 10^{-3} \text{ M}^{-1} \text{ minutes}^{-1}$ for H₃O⁺ catalysed aldol

216 reaction (Sareen et al., 2010). Naturally, both of these ions are present in the aqueous filter extract. Despite the
217 oligomerisation of methylglyoxal monomers during the 9 hours at room temperature, the derivatisation
218 proceeds as well during this time leading to higher methylglyoxal concentrations in sample b than in sample
219 a, which was directly measured after 15 hours heating (Fig. 1a).

220 Longer heating times than 24 hours (30 and 48 hours) led to lower methylglyoxal concentrations in
221 the filter samples as well. A possible explanation might be the decomposition of the derivatised compound
222 during the long heating process. As no decrease of the concentration was observed in the methylglyoxal
223 standard solution (Fig. 1b) the loss of the derivatisation group is unlikely as a reason for the lower
224 concentrations. Thus, it can be speculated that the low methylglyoxal concentrations in the filter samples are
225 a result of further reactions with particle-phase species, which do not exist in the standard samples. Based on
226 the outlined results, a heating time of 24 hours was chosen.

227 The pH during the heating process was investigated as well.

228

229 *Influence of pH*

230 The effect of the pH was examined with PTFE filters, which were sampled after OH-radical oxidation of 1,3,5-
231 TMB at RH = 50% in the presence of NH_4HSO_4 particles (experiment #2). The pH was varied between pH =
232 1 and pH = 7.

233 As it can be seen in Fig. 2a the highest methylglyoxal concentration can be found at pH = 1. The methylglyoxal
234 concentration was about two times higher at pH = 1 ($c = 1.01 \pm 0.11 \mu\text{mol L}^{-1}$) compared to the filter extract,
235 which was neither heated nor acidified ($c = 0.45 \pm 0.01 \mu\text{mol L}^{-1}$). An increasing pH leads to a lower
236 methylglyoxal concentration, which can be observed for filter samples (Fig. 2a) as well as for the
237 methylglyoxal standard solution (Fig. 2b). As this was observed for both types of samples it appears that the
238 pH influences the derivatisation and/or the oligomer decomposition. No influence of the pH on the PFBHA
239 derivatisation reaction was reported by Rodigast et al. (2015) indicating the effect of the pH is connected to
240 thermal decomposition of the heat-decomposable oligomeric compounds. In summary, based on these results
241 pH = 1 was used.

242

243 *Influence of heating temperature*

244 The effect of the heating temperature was examined with filter samples of experiment #2. The heating
245 temperature was varied between 50°C and 100°C and the filter extracts were heated for 24 hours at pH = 1. A
246 temperature above 100°C cannot be used to avoid evaporation of water and/or target compounds. Fig. 3a shows
247 the influence of the temperature on the decomposition of the heat-decomposable methylglyoxal oligomers into
248 monomers. Higher concentration of methylglyoxal can be detected with higher temperature. The results
249 illustrated in Fig. 3a indicate that a higher temperature than 50°C is needed to decompose the oligomeric
250 compounds. In comparison to the filter, which was neither acidified nor heated, the concentration increased by
251 a factor of two if the extract was heated to 100°C. Fig. 3b shows no significant influence of the temperature
252 on the methylglyoxal standard solution. Thus, an influence of the heating temperature on the derivatisation
253 procedure can be excluded.

254 Based on these results, the PTFE filter extracts from the aerosol chamber experiments were acidified
255 to pH = 1 and heated for 24 hours to 100°C to decompose heat-decomposable oligomeric compounds into

256 methylglyoxal. According to the literature studies other carbonyl compounds can be expected as particle-phase
257 products, e.g., propionaldehyde (Cocker et al., 2001), glyoxal (Cocker et al., 2001; Huang et al., 2015), 2-
258 methyl-4-oxo-2-pentenal (Healy et al., 2008; Huang et al., 2014), glycolaldehyde (Cocker et al., 2001) and
259 3,5-dimethylbenzaldehyde (Huang et al., 2014). Noticeably, no carbonyl compounds other than methylglyoxal
260 were identified, which showed an increase after thermal decomposition.

261 The developed quantification method was afterwards applied to laboratory-generated SOA formed
262 during further oxidation experiments of 1,3,5-TMB to investigate the influence of seed particle acidity and
263 relative humidity on the oligomer content.

264

265 3.2 SOA yield and growth curves of 1,3,5-TMB oxidation

266 SOA formation of 1,3,5-TMB was investigated in a number of literature studies mostly in the presence of NO_x
267 and under variation of the hydrocarbon to NO_x ratio ($[\text{HC}]/[\text{NO}_x]$ ratio). Healy et al. (2008) determined SOA
268 yields (Y_{SOA}) of 1,3,5-TMB photooxidation ranging from 4.5 to 8.3%. Further studies determined Y_{SOA}
269 between 0.29% and 15.6% (Table 4) under variation of the $[\text{HC}]/[\text{NO}_x]$ ratio concluding SOA formation is
270 enhanced at low NO_x mixing ratios (Sato et al., 2012; Kleindienst et al., 1999; Baltensperger et al., 2005;
271 Wyche et al., 2009; Odum et al., 1997; Paulsen et al., 2005; Cocker et al., 2001). Only Cao and Jang (2007)
272 investigated SOA yields in the absence of NO_x and reported values between 7.1 and 13.8%. The SOA yields
273 were also determined in the present study for all conducted experiments based on the ratio of ΔM to ΔHC
274 (Table 2). **Notably, the SOA yields of 1,3,5-TMB are not corrected for the wall loss to the surface of the aerosol**
275 **chamber, which might possibly lead to an underestimation of the reported SOA yields.** Y_{SOA} varied between 4
276 and 7% dependent on reaction conditions and is in good agreement to literature values.

277 For a further investigation of SOA-formation processes of 1,3,5-TMB, Fig. 4a illustrates the
278 dependency between the consumption of 1,3,5-TMB (ΔHC) and the produced organic particle mass (ΔM).
279 Particle growth started directly after the experiment was initialised indicating that the oxidation leads
280 immediately to the formation of condensable products as first-generation oxidation products. These products
281 condense on the pre-existing seed particles resulting in the immediate particle growth observed in Fig. 4a.
282 Differences of the growth curves in dependence on the seed particles (NH_4HSO_4 and $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$) were
283 not observed concluding that the seed particle acidity (Table 5) has no influence on the SOA formation of
284 1,3,5-TMB. Cao and Jang (2007) found also only a small influence of seed particle acidity on SOA formation.

285 Fig. 4a showed great differences in the growth curves under variation of RH. **The RH value can have**
286 **an influence on the phase state of the particles which, in turn, has an effect on the partitioning of the compounds**
287 **into the particles as well as on particle-phase reactions (Ziemann, 2010; Saukko et al., 2012). The**
288 **deliquescence RH of $(\text{NH}_4)_2\text{SO}_4$ and NH_4HSO_4 seed particles are known from literature to be 79% and 39%,**
289 **respectively (Cziczo et al., 1997). Thus, pure NH_4HSO_4 seed particles are liquid at the applied RH values of**
290 **50% and 75% and solid at $\text{RH} = 0\%$ while $(\text{NH}_4)_2\text{SO}_4$ seed particles are solid over the whole RH range.**
291 **Nevertheless, in the present study the particles are a mixture of inorganic and partitioned organic compounds.**
292 **Organic compounds might change the phase state of the particles (Bertram et al., 2011) due to an influence on**
293 **the deliquescence point (Andrews and Larson, 1993; Lightstone et al., 2000) as well as the hygroscopic**
294 **behaviour of the particles (Lightstone et al., 2000; Prenni et al., 2003; Chen and Lee, 1999). Virtanen et al.,**
295 **2010 postulated that particles are in an amorphous solid state if oligomeric compounds are present in the**

296 particles. Thus, it can be speculated that the particles in the present study containing a fraction of up to 8% of
297 oligomeric compounds might be in an amorphous solid phase state. It was assumed, that the further reactions
298 in the particle phase might be inhibited in solid particles (Saukko et al., 2012) thus, further oligomerisation or
299 other reactions might become less effective after a certain fraction of oligomers exist in the particles. In
300 addition, the partitioning of methylglyoxal monomers can be inhibited into solid particles (Saukko et al., 2012),
301 which then might also lead to lower oligomer fractions in SOA. Thus, it might be possible that the phase state
302 influences ΔM and the SOA yields. ΔM is the highest at RH = 0% ($\Delta M = 18.1 - 19.7 \mu\text{g m}^{-3}$) whereas ΔM is
303 the lowest under humid conditions (RH = 50% and 75%, $\Delta M = 11.3 - 11.7 \mu\text{g m}^{-3}$ and $13.9 - 14.2 \mu\text{g m}^{-3}$).
304 Due to the variation of RH in the aerosol chamber the liquid water content (LWC) of the particles is changing
305 (Table 5). The LWC was calculated using model II from the extended aerosol thermodynamic model (E-AIM;
306 Clegg et al., 1998). With increasing RH the LWC of the seed particle increases as well. The LWC of the seed
307 particles influences i) the partitioning of the compounds from the gas phase into the particle phase and ii) the
308 formation and/or further reaction in the particle phase (Zuend et al., 2010; Cocker et al., 2001; Seinfeld et al.,
309 2001; Fick et al., 2003). These two effects might influence the SOA formation under different relative
310 humidities.

311 An effect can also be seen in Fig. 4b. The SOA formation is enhanced at RH = 0% leading to the
312 highest SOA yields of $Y_{\text{SOA}} \approx 7\%$ for both seed particles. Higher RH values resulted in lower Y_{SOA} between 4
313 and 5%. These findings are in good agreement with the study by Cao and Jang (2007), which observed lower
314 Y_{SOA} values at elevated RH.

315 The influence of RH on SOA formation is controversial (Hennigan et al., 2008; Fick et al., 2003;
316 Edney et al., 2000; Saxena and Hildemann, 1996; Baker et al., 2001; Hasson et al., 2001, Cocker et al., 2001).
317 Edney et al. (2000) and Seinfeld et al. (2001) reported an enhanced SOA formation of hydrophilic compounds
318 under humid conditions and a lowered SOA formation of hydrophobic compounds. This is also supported by
319 Saxena and Hildemann (1996), which found an enhanced partitioning of organic compounds with several
320 hydroxyl groups at higher LWCs of the particles. This might lead to the conclusion that the OH-radical
321 oxidation of 1,3,5-TMB results in the formation of hydrophobic compounds which showed an enhanced
322 partitioning under dry conditions. Additionally, it can be speculated that the formation of oligomeric
323 compounds can be enhanced at lower RH values resulting in higher Y_{SOA} due to the increasing conversion of
324 the monomeric building blocks and their enhanced partitioning into the particle phase.

325

326 3.3 Particulate methylglyoxal

327 Methylglyoxal is reported in the literature as an important oxidation product of 1,3,5-TMB (Metzger et al.,
328 2008; Healy et al., 2008; Cocker et al., 2001; Smith et al., 1999; Wyche et al., 2009; Baltensperger et al., 2005;
329 Rickard et al., 2010; Kalberer et al., 2004; Yu et al., 1997; Kleindienst et al., 1999; Müller et al., 2012; Nishino
330 et al., 2010; Hamilton et al., 2003; Tuazon et al., 1986; Bandow and Washida, 1985) with yields in the particle
331 phase between 0.7 and 2%.

332 The fraction of methylglyoxal in the particle phase in dependency on the reaction conditions is shown
333 in Fig. 5a, with resulting fractions between $\approx 0.6\%$ and $\approx 2.2\%$. With increasing RH the fraction decreases for
334 both seed particles. Methylglyoxal has the highest fraction under dry conditions ($1.73 \pm 0.20\%$ and
335 $2.17 \pm 0.20\%$ for NH_4HSO_4 and $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$) and with $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles.

336 Healy et al. (2008) measured a contribution of methylglyoxal to SOA mass of $2.06 \pm 0.08\%$ from the
337 photooxidation of 1,3,5-TMB in the presence of NO_x at $\text{RH} = 50\%$. In the present study a methylglyoxal
338 fraction of $1.24 \pm 0.04\%$ for NH_4HSO_4 seed particles and $0.80 \pm 0.08\%$ for $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles
339 was determined at $\text{RH} = 50\%$. Thus, the contribution is slightly lower than measured by Healy et al. (2008).
340 In comparison, Cocker et al., 2001 measured a particulate fraction of 0.72% .

341 The dependency of particulate methylglyoxal on RH could be a result of the influence of RH on the
342 partitioning from the gas- into the particle phase or on further reactions in the particle phase forming oligomers.
343 The formation of oligomeric compounds from methylglyoxal has been investigated in a number of studies
344 (e. g. De Haan et al., 2009; Kalberer et al., 2004; Loeffler et al., 2006; Zhao et al., 2006; Sareen et al., 2010;
345 Altieri et al., 2008).

346

347 **3.4 Heat-decomposable methylglyoxal oligomers**

348 A method was developed to determine the contribution of heat-decomposable methylglyoxal oligomers to the
349 produced organic particle mass ΔM . The method is based on the thermal decomposition of the heat-
350 decomposable methylglyoxal oligomers into monomers. Thus, the concentration of monomeric methylglyoxal
351 was determined prior and after thermal decomposition. The concentrations were converted into the fraction of
352 methylglyoxal oligomers of ΔM using the molar mass of methylglyoxal ($M_w = 72.06 \text{ g mol}^{-1}$). An oligomer
353 fraction of ≈ 2 up to $\approx 8\%$ was determined.

354 Fig. 5b shows the dependency of the detected heat-decomposable methylglyoxal oligomers on the
355 relative humidity with NH_4HSO_4 and $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles. In the presence of NH_4HSO_4 seed
356 particles the highest oligomer fraction ($8.2 \pm 0.7\%$) can be observed with $\text{RH} = 0\%$ whereas in the presence of
357 $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles the oligomer fraction is the lowest ($2.1 \pm 0.4\%$) under dry conditions
358 ($\text{RH} = 0\%$). A possible explanation for the opposite trend of the oligomer fractions with RH between NH_4HSO_4
359 and $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles could be different oligomer formation mechanism caused by different
360 seed particle acidity. The type of accretion reaction might change with pH (Yasmeen et al., 2010). In Table 5
361 the pH of the seed particles was calculated with E-AIM. NH_4HSO_4 particles have $\text{pH} = 0.1$ and 1.2 at
362 $\text{RH} = 50\%$ and 75% . In comparison $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles are less acidic ($\text{pH} = 4.0$ at $\text{RH} = 50\%$ and
363 $\text{pH} = 4.2$ at $\text{RH} = 75\%$).

364 It was postulated by Yasmeen et al. (2010) that a lower pH ($\text{pH} < 3.5$) favor acetal/hemiacetal
365 formation whereas at high pH ($\text{pH} = 4 - 5$) aldol condensation are more relevant. This has been supported by
366 Sedehi et al. (2013) and Sareen et al. (2010).

367 Thus, in the presence of strong acidic NH_4HSO_4 seed particles acetal/hemiacetal formation might be
368 the favored oligomer formation mechanism. Oligomerisation via acetal/hemiacetal formation occurs under a
369 reversible water loss (Yasmeen et al., 2010). As higher RH values in the aerosol chamber LEAK leads to higher
370 LWCs of the seed particles (Table 5) the chemical equilibrium of the reaction shifts towards the precursor
371 compound resulting in a lower methylglyoxal oligomer fraction (Kalberer et al., 2004; Liggió et al., 2005). In
372 addition, the pH of NH_4HSO_4 particles decreases with decreasing RH (Table 5) thus acid-catalysed
373 acetal/hemiacetal formation might be enhanced under dry conditions due to a lower pH (Liggió et al., 2005).

374 In the presence of $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles the oligomer fraction increases with increasing
375 RH (Fig. 5b). As it was mentioned, aldol condensation can be assumed as the favored accretion reaction under

376 these conditions (Yasmeen et al., 2010). Aldol condensation includes as a first step aldol addition followed by
377 a loss of water. The loss of water is irreversible, thus the aldol condensation will not be inhibited with higher
378 LWC of the seed particles.

379 Other accretion reactions can contribute to the formation of heat-decomposable methylglyoxal
380 oligomers with $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles as well. Altieri et al. (2008) detected products formed through
381 acid-catalysed esterification at $\text{pH} \approx 4$. This equilibrium reaction involves the reversible loss of water as it was
382 reported for acetal/hemiacetal formation (Lim et al., 2010). Thus, it can be expected that with higher LWCs
383 the contribution of esterification reactions to oligomer formation decreases due to the shift of the equilibrium
384 towards the monomers.

385 Imidazole formation was also postulated as possible oligomer-formation mechanism for
386 methylglyoxal (Sedehehi et al., 2013; De Haan et al., 2011). It was found that imidazole formation is of minor
387 importance compared to aldol condensation (Sedehehi et al., 2013). However, imidazole formation involves also
388 a reversible loss of water, thus it does not provide a feasible explanation for the higher oligomer fraction at
389 higher RH with $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles.

390 Radical – radical reactions are also postulated as a possible reaction pathway to form oligomers
391 (Schaefer et al., 2015; Lim et al., 2013; Rincon et al., 2009; Lim et al., 2010; Sun et al., 2010). Radical-radical
392 reactions of methylglyoxal might occur following the H - atom abstraction of methylglyoxal with OH radicals
393 and a subsequent recombination of the resulting alkyl radicals (as discussed for glyoxal in Schaefer et al.,
394 2015). The contribution of radical – radical reactions to oligomer formation is not well understood as
395 obviously, the reaction of alkyl radicals with oxygen tends to suppress this pathway. Nevertheless, it can be
396 expected, that with higher LWCs of the seed particles and thus with a higher reaction volume, the absolute
397 amount of methylglyoxal in the particle phase might increase, but not its particle-phase concentrations. For
398 aerosol particle systems, ionic strength effects (Herrmann et al., 2015) are able to influence the uptake of
399 methylglyoxal into the particle phase as well. Waxman et al. (2015) observed a salting-out effect for
400 methylglyoxal for all investigated seed particles at higher ionic strengths.

401 Low pH combined with high solute concentrations as calculated for the present aerosol particles, can
402 trigger isomerisation (or switching) reactions as discussed by Herrmann et al. (2015). Overall, a clear
403 discussion on how radical-radical reaction might be affected by increasing LWC (through increasing RH) and
404 by pH is difficult at the current level of knowledge.

405 In summary, the present study provides a reliable quantification method for heat-decomposable
406 methylglyoxal oligomers formed by 1,3,5-TMB oxidation. The fraction of oligomeric substances formed solely
407 by methylglyoxal oligomerisation varied dependent on RH and seed particle acidity between 2 – 8%, which is
408 lower than the determined values by Kalberer et al. (2004) and Baltensperger et al. (2005) (varying between
409 50 and 80%). **In the present study only heat-decomposable methylglyoxal oligomers were quantified, thus
410 there might be not heat-decomposable methylglyoxal oligomers or oligomers originating from other monomers
411 than methylglyoxal, which were not determined in the present study leading to lower oligomer fractions of ΔM
412 compared to the literature studies.** The obtained data are not fully conclusive and literature studies are often
413 contradicting. Thus, more experiments are necessary to get a clearer picture about the influence of RH and
414 particle phase acidity on oligomer formation and to explain the non-linear relation between the oligomer
415 fraction and RH.

416

417 **4. Summary**

418 In the present study a method was developed to quantify oligomers formed from methylglyoxal. The method
419 is based on the thermal decomposition of heat-decomposable methylglyoxal oligomers into monomers. The
420 formed methylglyoxal monomers were detected with PFBHA derivatisation and GC/MS analysis. The
421 influence of heating time, pH and heating temperature on the decomposition of heat-decomposable
422 methylglyoxal oligomers was systematically investigated. The best result was achieved with a heating time of
423 24 hours at 100°C and pH = 1. The method was applied to heat-decomposable methylglyoxal oligomers
424 formed during the oxidation of 1,3,5-TMB resulting in an oligomer fraction of up to ≈ 8%. A contradicting
425 dependency of the oligomer fraction under varying RH with NH₄HSO₄ and (NH₄)₂SO₄/H₂SO₄ seed particles
426 was found, which might be caused by different oligomer formation mechanisms.
427 Overall, the present method provides an important step revealing the amount of oligomers present in the
428 particle phase, their tentative formation mechanism and their importance for aqSOA formation.

429

430 **5. Acknowledgements**

431 This study was supported by the Scholarship program of the German Federal Environmental Foundation
432 (Deutsche Bundesstiftung Umwelt, DBU; grant number 20013/244).

433

434 **References**

- 435 Alfarra, M. R., Paulsen, D., Gysel, M., Garforth, A. A., Dommen, J., Prevot, A. S. H., Worsnop, D. R.,
436 Baltensperger, U., and Coe, H.: A mass spectrometric study of secondary organic aerosols formed from the
437 photooxidation of anthropogenic and biogenic precursors in a reaction chamber, *Atmospheric Chemistry and*
438 *Physics*, 6, 5279-5293, 10.5194/acp-6-5279-2006, 2006.
- 439 Altieri, K. E., Seitzinger, S. P., Carlton, A. G., Turpin, B. J., Klein, G. C., and Marshall, A. G.: Oligomers
440 formed through in-cloud methylglyoxal reactions: Chemical composition, properties, and mechanisms
441 investigated by ultra-high resolution FT-ICR mass spectrometry, *Atmospheric Environment*, 42, 1476-1490,
442 10.1016/j.atmosenv.2007.11.015, 2008.
- 443 **Andrews, E., and Larson, S. M.: Effect of surfactant layers on the size changes of aerosol-particles as a function**
444 **of relative-humidity, *Environmental Science & Technology*, 27, 857-865, 10.1021/es00042a007, 1993.**
- 445 Angove, D. E., Fookes, C. J. R., Hynes, R. G., Walters, C. K., and Azzi, M.: The characterisation of secondary
446 organic aerosol formed during the photodecomposition of 1,3-butadiene in air containing nitric oxide,
447 *Atmospheric Environment*, 40, 4597-4607, 10.1016/j.atmosenv.2006.03.046, 2006.
- 448 Bahreini, R., Keywood, M. D., Ng, N. L., Varutbangkul, V., Gao, S., Flagan, R. C., Seinfeld, J. H., Worsnop,
449 D. R., and Jimenez, J. L.: Measurements of secondary organic aerosol from oxidation of cycloalkenes,
450 terpenes, and m-xylene using an Aerodyne aerosol mass spectrometer, *Environmental Science &*
451 *Technology*, 39, 5674-5688, 10.1021/es048061a, 2005.
- 452 Baker, J., Aschmann, S. M., Arey, J., and Atkinson, R.: Reactions of stabilized Criegee intermediates from the
453 gas-phase reactions of O-3 with selected alkenes, *International Journal of Chemical Kinetics*, 34, 73-85,
454 2001.
- 455 Baltensperger, U., Kalberer, M., Dommen, J., Paulsen, D., Alfarra, M. R., Coe, H., Fisseha, R., Gascho, A.,
456 Gysel, M., Nyeki, S., Sax, M., Steinbacher, M., Prevot, A. S. H., Sjogren, S., Weingartner, E., and Zenobi,
457 R.: Secondary organic aerosols from anthropogenic and biogenic precursors, *Faraday Discussions*, 130, 265-
458 278, 10.1039/b417367h, 2005.
- 459 Bandow, H., and Washida, N.: Ring-cleavage reactions of aromatic hydrocarbons studied by FT-IR
460 spectroscopy .III. Photooxidation of 1,2,3-,1,2,4-, and 1,3,5-trimethylbenzene in the NO_x-air system,
461 *Bulletin of the Chemical Society of Japan*, 58, 2549-2555, 10.1246/bcsj.58.2549, 1985.
- 462 Barsanti, K. C., and Pankow, J. F.: Thermodynamics of the formation of atmospheric organic particulate matter
463 by accretion reactions - 2. Dialdehydes, methylglyoxal, and diketones, *Atmospheric Environment*, 39, 6597-
464 6607, 10.1016/j.atmosenv.2005.07.056, 2005.
- 465 Berndt, T., and Böge, O.: Formation of phenol and carbonyls from the atmospheric reaction of OH radicals
466 with benzene, *Physical Chemistry Chemical Physics*, 8, 1205-1214, 10.1039/b514148f, 2006.

467 Bertram, A. K., Martin, S. T., Hanna, S. J., Smith, M. L., Bodsworth, A., Chen, Q., Kuwata, M., Liu, A.; You,
468 Y., and Zorn, S. R.: Predicting the relative humidities of liquid-liquid phase separation, efflorescence, and
469 deliquescence of mixed particles of ammonium sulfate, organic material, and water using the organic-to-
470 sulfate mass ratio of the particle and the oxygen-to-carbon elemental ratio of the organic component,
471 *Atmospheric Chemistry and Physics*, 11, 10995-11006, 10.5194/acp-11-10995-2011, 2011.

472 Birdsall, A. W., Zentner, C. A., and Elrod, M. J.: Study of the kinetics and equilibria of the oligomerization
473 reactions of 2-methylglyceric acid, *Atmospheric Chemistry and Physics*, 13, 3097-3109, 10.5194/acp-13-
474 3097-2013, 2013.

475 Bones, D. L., Henricksen, D. K., Mang, S. A., Gonsior, M., Bateman, A. P., Nguyen, T. B., Cooper, W. J., and
476 Nizkorodov, S. A.: Appearance of strong absorbers and fluorophores in limonene-O-3 secondary organic
477 aerosol due to NH₄⁺-mediated chemical aging over long time scales, *Journal of Geophysical Research-
478 Atmospheres*, 115, 14, 10.1029/2009jd012864, 2010.

479 Cao, G., and Jang, M.: Effects of particle acidity and UV light on secondary organic aerosol formation from
480 oxidation of aromatics in the absence of NO_x, *Atmospheric Environment*, 41, 7603-7613,
481 10.1016/j.atmosenv.2007.05.034, 2007.

482 Casale, M. T., Richman, A. R., Elrod, M. J., Garland, R. M., Beaver, M. R., and Tolbert, M. A.: Kinetics of
483 acid-catalyzed aldol condensation reactions of aliphatic aldehydes, *Atmospheric Environment*, 41, 6212-
484 6224, 10.1016/j.atmosenv.2007.04.002, 2007.

485 Chen, Y. Y., and Lee, W. M. G.: Hygroscopic properties of inorganic-salt aerosol with surface-active organic
486 compounds, *Chemosphere*, 38, 2431-2448, 10.1016/s0045-6535(98)00436-6, 1999.

487 Clegg, S. L., Brimblecombe, P., and Wexler, A. S.: Thermodynamic model of the system H⁺-NH₄⁺-SO₄²⁻-
488 -NO₃⁻-H₂O at tropospheric temperatures, *Journal of physical Chemistry A*, 102, 2137-2154,
489 10.1021/jp973042r, 1998.

490 Cocker, D. R., Mader, B. T., Kalberer, M., Flagan, R. C., and Seinfeld, J. H.: The effect of water on gas-particle
491 partitioning of secondary organic aerosol: II. m-xylene and 1,3,5-trimethylbenzene photooxidation systems,
492 *Atmospheric Environment*, 35, 6073-6085, 10.1016/s1352-2310(01)00405-8, 2001.

493 Cziczo, D. J., Nowak, J. B., Hu, J. H., and Abbatt, J. P. D.: Infrared spectroscopy of model tropospheric
494 aerosols as a function of relative humidity: Observation of deliquescence and crystallization, *Journal of
495 Geophysical Research-Atmospheres*, 102, 18843-18850, 10.1029/97jd01361, 1997.

496 De Haan, D. O., Corrigan, A. L., Tolbert, M. A., Jimenez, J. L., Wood, S. E., and Turley, J. J.: Secondary
497 organic aerosol formation by self-reactions of methylglyoxal and glyoxal in evaporating droplets,
498 *Environmental Science & Technology*, 43, 8184-8190, 10.1021/es902152t, 2009.

499 De Haan, D. O., Hawkins, L. N., Kononenko, J. A., Turley, J. J., Corrigan, A. L., Tolbert, M. A., and Jimenez,
500 J. L.: Formation of Nitrogen-Containing Oligomers by Methylglyoxal and Amines in Simulated Evaporating
501 Cloud Droplets, *Environmental Science & Technology*, 45, 984-991, 10.1021/es102933x, 2011.

502 Denkenberger, K. A., Moffet, R. C., Holecek, J. C., Rebotier, T. P., and Prather, K. A.: Real-time, single-
503 particle measurements of oligomers in aged ambient aerosol particles, *Environmental Science & Technology*,
504 41, 5439-5446, 10.1021/es0703291, 2007.

505 Dommen, J., Metzger, A., Duplissy, J., Kalberer, M., Alfarra, M. R., Gascho, A., Weingartner, E., Prevot, A.
506 S. H., Verheggen, B., and Baltensperger, U.: Laboratory observation of oligomers in the aerosol from
507 isoprene/NO_x photooxidation, *Geophysical Research Letters*, 33, 1-5, 10.1029/2006GL026523, 2006.

508 Drozd, G. T., and McNeill, V. F.: Organic matrix effects on the formation of light-absorbing compounds from
509 alpha-dicarbonyls in aqueous salt solution, *Environmental Science-Processes & Impacts*, 16, 741-747,
510 10.1039/c3em00579h, 2014.

511 Edney, E. O., Driscoll, D. J., Speer, R. E., Weathers, W. S., Kleindienst, T. E., Li, W., and Smith, D. F.: Impact
512 of aerosol liquid water on secondary organic aerosol yields of irradiated toluene/propylene/NO_x/(NH₄)₂
513 SO₄/air mixtures, *Atmospheric Environment*, 34, 3907-3919, 10.1016/s1352-2310(00)00174-6, 2000.

514 Fick, J., Pommer, L., Nilsson, C., and Andersson, B.: Effect of OH radicals, relative humidity, and time on the
515 composition of the products formed in the ozonolysis of alpha-pinene, *Atmospheric Environment*, 37, 4087-
516 4096, 10.1016/s1352-2310(03)00522-3, 2003.

517 Gao, S., Keywood, M., Ng, N. L., Surratt, J., Varutbangkul, V., Bahreini, R., Flagan, R. C., and Seinfeld, J.
518 H.: Low-molecular-weight and oligomeric components in secondary organic aerosol from the ozonolysis of
519 cycloalkenes and alpha-pinene, *Journal of Physical Chemistry A*, 108, 10147-10164, 10.1021/jp047466e,
520 2004.

521 Garland, R. M., Elrod, M. J., Kincaid, K., Beaver, M. R., Jimenez, J. L., and Tolbert, M. A.: Acid-catalyzed
522 reactions of hexanal on sulfuric acid particles: Identification of reaction products, *Atmospheric Environment*,
523 40, 6863-6878, 10.1016/j.atmosenv.2006.07.009, 2006.

524 Gee, I. L., and Sollars, C. J.: Ambient air levels of volatile organic compounds in Latin American and Asian
525 cities, *Chemosphere*, 36, 2497-2506, 10.1016/S0045-6535(97)10217-X, 1998.

526 Gutbrod, R., Schindler, R. N., Kraka, E., and Cremer, D.: Formation of OH radicals in the gas phase ozonolysis
527 of alkenes: the unexpected role of carbonyl oxides, *Chemical Physics Letters*, 252, 221-229, 10.1016/0009-
528 2614(96)00126-1, 1996.

529 Hall, W. A., and Johnston, M. V.: Oligomer Formation Pathways in Secondary Organic Aerosol from MS and
530 MS/MS Measurements with High Mass Accuracy and Resolving Power, *Journal of the American Society
531 for Mass Spectrometry*, 23, 1097-1108, 10.1007/s13361-012-0362-6, 2012.

532 Hallquist, M., Wenger, J. C., Baltensperger, U., Rudich, Y., Simpson, D., Claeys, M., Dommen, J., Donahue,
533 N. M., George, C., Goldstein, A. H., Hamilton, J. F., Herrmann, H., Hoffmann, T., Iinuma, Y., Jang, M.,
534 Jenkin, M. E., Jimenez, J. L., Kiendler-Scharr, A., Maenhaut, W., McFiggans, G., Mentel, T. F., Monod, A.,
535 Prevot, A. S. H., Seinfeld, J. H., Surratt, J. D., Szmigielski, R., and Wildt, J.: The formation, properties and
536 impact of secondary organic aerosol: current and emerging issues, *Atmospheric Chemistry and Physics*, 9,
537 5155-5236, 10.5194/acp-9-5155-2009, 2009.

538 Hamilton, J. F., Lewis, A. C., Bloss, C., Wagner, V., Henderson, A. P., Golding, B. T., Wirtz, K., Martin-
539 Reviejo, M., and Pilling, M. J.: Measurements of photo-oxidation products from the reaction of a series of
540 alkyl-benzenes with hydroxyl radicals during EXACT using comprehensive gas chromatography,
541 *Atmospheric Chemistry and Physics*, 3, 1999-2014, 10.5194/acp-3-1999-2003, 2003.

542 Hamilton, J. F., Lewis, A. C., Reynolds, J. C., Carpenter, L. J., and Lubben, A.: Investigating the composition
543 of organic aerosol resulting from cyclohexene ozonolysis: low molecular weight and heterogeneous reaction
544 products, *Atmospheric Chemistry and Physics*, 6, 4973-4984, 2006.

545 Hasson, A. S., Orzechowska, G., and Paulson, S. E.: Production of stabilized Criegee intermediates and
546 peroxides in the gas phase ozonolysis of alkenes 1. Ethene, trans-2-butene, and 2,3-dimethyl-2-butene,
547 *Journal of Geophysical Research-Atmospheres*, 106, 34131-34142, 10.1029/2001jd000597, 2001.

548 Hastings, W. P., Koehler, C. A., Bailey, E. L., and De Haan, D. O.: Secondary organic aerosol formation by
549 glyoxal hydration and oligomer formation: Humidity effects and equilibrium shifts during analysis,
550 *Environmental Science & Technology*, 39, 8728-8735, 10.1021/es050446l, 2005.

551 Healy, R. M., Wenger, J. C., Metzger, A., Duplissy, J., Kalberer, M., and Dommen, J.: Gas/particle partitioning
552 of carbonyls in the photooxidation of isoprene and 1,3,5-trimethylbenzene, *Atmospheric Chemistry and
553 Physics*, 8, 3215-3230, 10.5194/acp-8-3215-2008, 2008.

554 Heaton, K. J., Dreyfus, M. A., Wang, S., and Johnston, M. V.: Oligomers in the early stage of biogenic
555 secondary organic aerosol formation and growth, *Environmental Science & Technology*, 41, 6129-6136,
556 10.1021/es070314n, 2007.

557 Hennigan, C. J., Bergin, M. H., Dibb, J. E., and Weber, R. J.: Enhanced secondary organic aerosol formation
558 due to water uptake by fine particles, *Geophysical Research Letters*, 35, 5, 10.1029/2008gl035046, 2008.

559 Herrmann, H., Schaefer, T., Tilgner, A., Styler, S. A., Weller, C., Teich, M., and Otto, T.: Tropospheric
560 aqueous-phase chemistry: Kinetics, mechanisms, and its coupling to a changing gas phase, *Chemical
561 Reviews*, 115, 4259-4334, 10.1021/cr500447k, 2015.

562 Holmes, B. J., and Petrucci, G. A.: Water-soluble oligomer formation from acid-catalyzed reactions of
563 levoglucosan in proxies of atmospheric aqueous aerosols, *Environmental Science & Technology*, 40, 4983-
564 4989, 10.1021/es060646c, 2006.

565 Huang, M. Q., Lin, Y. H., Huang, X. Y., Liu, X. Q., Guo, X. Y., Hu, C. J., Zhao, W. X., Gu, X. J., Fang, L.,
566 and Zhang, W. J.: Experimental study of particulate products for aging of 1,3,5-trimethylbenzene secondary
567 organic aerosol, *Atmospheric Pollution Research*, 6, 209-219, 10.5094/APR.2015.025, 2015.

568 Iinuma, Y., Böge, O., Gnauk, T., and Herrmann, H.: Aerosol-chamber study of the α -pinene/O₃ reaction:
569 influence of particle acidity on aerosol yields and products, *Atmospheric Environment*, 38, 761-773,
570 10.1016/j.atmosenv.2003.10.015, 2004.

571 Jang, M., Lee, S., and Kamens, R. M.: Organic aerosol growth by acid-catalyzed heterogeneous reactions of
572 octanal in a flow reactor, *Atmospheric Environment*, 37, 2125-2138, 10.1016/s1352-2310(03)00077-3, 2003.

573 Jang, M. S., and Kamens, R. M.: Atmospheric secondary aerosol formation by heterogeneous reactions of
574 aldehydes in the presence of a sulfuric acid aerosol catalyst, *Environmental Science & Technology*, 35, 4758-
575 4766, 10.1021/es010790s, 2001.

576 Kahnt, A., Iinuma, Y., Böge, O., Mutzel, A., and Herrmann, H.: Denuder sampling techniques for the
577 determination of gas-phase carbonyl compounds: A comparison and characterisation of in situ and ex situ
578 derivatisation methods, *Journal of Chromatography B*, 879, 1402-1411, 10.1016/j.jchromb.2011.02.028,
579 2011.

580 Kalberer, M., Paulsen, D., Sax, M., Steinbacher, M., Dommen, J., Prevot, A. S. H., Fisseha, R., Weingartner,
581 E., Frankevich, V., Zenobi, R., and Baltensperger, U.: Identification of polymers as major components of
582 atmospheric organic aerosols, *Science*, 303, 1659-1662, 10.1126/science.1092185, 2004.

583 Kalberer, M., Sax, M., and Samburova, V.: Molecular size evolution of oligomers in organic aerosols collected
584 in urban atmospheres and generated in a smog chamber, *Environmental Science & Technology*, 40, 5917-
585 5922, 10.1021/es052576o, 2006.

586 Khoder, M. I.: Ambient levels of volatile organic compounds in the atmosphere of Greater Cairo, *Atmospheric*
587 *Environment*, 41, 554-566, 10.1016/j.atmosenv.2006.08.051, 2007.

588 Kleindienst, T. E., Smith, D. F., Li, W., Edney, E. O., Driscoll, D. J., Speer, R. E., and Weathers, W. S.:
589 Secondary organic aerosol formation from the oxidation of aromatic hydrocarbons in the presence of dry
590 submicron ammonium sulfate aerosol, *Atmospheric Environment*, 33, 3669-3681, 10.1016/S1352-
591 2310(99)00121-1, 1999.

592 Krizner, H. E., De Haan, D. O., and Kua, J.: Thermodynamics and Kinetics of Methylglyoxal Dimer
593 Formation: A Computational Study, *Journal of Physical Chemistry A*, 113, 6994-7001, 10.1021/jp903213k,
594 2009.

595 Kua, J., Galloway, M. M., Millage, K. D., Avila, J. E., and De Haan, D. O.: Glycolaldehyde Monomer and
596 Oligomer Equilibria in Aqueous Solution: Comparing Computational Chemistry and NMR Data, *Journal of*
597 *Physical Chemistry A*, 117, 2997-3008, 10.1021/jp312202j, 2013.

598 Kundu, S., Fisseha, R., Putman, A. L., Rahn, T. A., and Mazzoleni, L. R.: High molecular weight SOA
599 formation during limonene ozonolysis: insights from ultrahigh-resolution FT-ICR mass spectrometry
600 characterization, *Atmospheric Chemistry and Physics*, 12, 5523-5536, 10.5194/acp-12-5523-2012, 2012.

601 Liggio, J., Li, S. M., and McLaren, R.: Heterogeneous reactions of glyoxal on particulate matter: Identification
602 of acetals and sulfate esters, *Environmental Science & Technology*, 39, 1532-1541, 10.1021/es048375y,
603 2005.

604 Lightstone, J. M., Onasch, T. B., Imre, D., and Oatis, S.: Deliquescence, efflorescence, and water activity in
605 ammonium nitrate and mixed ammonium nitrate/succinic acid microparticles, *Journal of Physical Chemistry*
606 *A*, 104, 9337-9346, 10.1021/jp002137h, 2000.

607 Lim, Y. B., Tan, Y., Perri, M. J., Seitzinger, S. P., and Turpin, B. J.: Aqueous chemistry and its role in
608 secondary organic aerosol (SOA) formation, *Atmospheric Chemistry and Physics*, 10, 10521-10539,
609 10.5194/acp-10-10521-2010, 2010.

610 Lim, Y. B., Tan, Y., and Turpin, B. J.: Chemical insights, explicit chemistry, and yields of secondary organic
611 aerosol from OH radical oxidation of methylglyoxal and glyoxal in the aqueous phase, *Atmospheric*
612 *Chemistry and Physics*, 13, 8651-8667, 10.5194/acp-13-8651-2013, 2013.

613 Lim, Y. B., and Turpin, B. J.: Laboratory evidence of organic peroxide and peroxyhemiacetal formation in the
614 aqueous phase and implications for aqueous OH, *Atmospheric Chemistry and Physics*, 15, 12867-12877,
615 10.5194/acp-15-12867-2015, 2015.

616 Liu, Y., Shao, M., Lu, S. H., Liao, C. C., Wang, J. L., and Chen, G.: Volatile organic compound (VOC)
617 measurements in the pearl river delta (PRD) region, China, *Atmospheric Chemistry and Physics*, 8, 1531-
618 1545, 2008.

619 Loeffler, K. W., Koehler, C. A., Paul, N. M., and De Haan, D. O.: Oligomer formation in evaporating aqueous
620 glyoxal and methyl glyoxal solutions, *Environmental Science & Technology*, 40, 6318-6323,
621 10.1021/es060810w, 2006.

622 Metzger, A., Dommen, J., Gaeggeler, K., Duplissy, J., Prevot, A. S. H., Kleffmann, J., Elshorbany, Y.,
623 Wisthaler, A., and Baltensperger, U.: Evaluation of 1,3,5 trimethylbenzene degradation in the detailed
624 tropospheric chemistry mechanism, MCMv3.1, using environmental chamber data, *Atmospheric Chemistry*
625 *and Physics*, 8, 6453-6468, 10.5194/acp-8-6453-2008, 2008.

626 Müller, L., Reinnig, M. C., Warnke, J., and Hoffmann, T.: Unambiguous identification of esters as oligomers
627 in secondary organic aerosol formed from cyclohexene and cyclohexene/alpha-pinene ozonolysis,
628 *Atmospheric Chemistry and Physics*, 8, 1423-1433, 2008.

629 Müller, M., Graus, M., Wisthaler, A., Hansel, A., Metzger, A., Dommen, J., and Baltensperger, U.: Analysis
630 of high mass resolution PTR-TOF mass spectra from 1,3,5-trimethylbenzene (TMB) environmental chamber
631 experiments, *Atmospheric Chemistry and Physics*, 12, 829-843, 10.5194/acp-12-829-2012, 2012.

632 Mutzel, A., Rodigast, M., Iinuma, Y., Boge, O., and Herrmann, H.: Monoterpene SOA - Contribution of first-
633 generation oxidation products to formation and chemical composition, *Atmospheric Environment*, 130, 136-
634 144, 10.1016/j.atmosenv.2015.10.080, 2016.

635 Nemet, I., Vikić-Topić, D., and Varga-Defterdarović, L.: Spectroscopic studies of methylglyoxal in water and
636 dimethylsulfoxide, *Bioorganic Chemistry*, 32, 560-570, 10.1016/j.bioorg.2004.05.008, 2004.

637 Nguyen, T. B., Roach, P. J., Laskin, J., Laskin, A., and Nizkorodov, S. A.: Effect of humidity on the
638 composition of isoprene photooxidation secondary organic aerosol, *Atmospheric Chemistry and Physics*, 11,
639 6931-6944, 10.5194/acp-11-6931-2011, 2011.

640 Nishino, N., Arey, J., and Atkinson, R.: Formation yields of glyoxal and methylglyoxal from the gas-phase
641 OH radical-initiated reactions of toluene, xylenes, and trimethylbenzenes as a function of NO₂ concentration,
642 *Journal of Physical Chemistry A*, 114, 10140-10147, 10.1021/jp105112h, 2010.

643 Nozière, B., and Esteve, W.: Organic reactions increasing the absorption index of atmospheric sulfuric acid
644 aerosols, *Geophysical Research Letters*, 32, 5, 10.1029/2004gl021942, 2005.

645 Noziere, B., Dziedzic, P., and Cordova, A.: Formation of secondary light-absorbing "fulvic-like" oligomers:
646 A common process in aqueous and ionic atmospheric particles?, *Geophysical Research Letters*, 34, 1-5,
647 10.1029/2007GL031300, 2007.

648 Noziere, B., and Cordova, A.: A kinetic and mechanistic study of the amino acid catalyzed aldol condensation
649 of acetaldehyde in aqueous and salt solutions, *Journal of Physical Chemistry A*, 112, 2827-2837,
650 10.1021/jp7096845, 2008.

651 Noziere, B., Dziedzic, P., and Cordova, A.: Inorganic ammonium salts and carbonate salts are efficient
652 catalysts for aldol condensation in atmospheric aerosols, *Physical Chemistry Chemical Physics*, 12, 3864-
653 3872, 10.1039/B924443C, 2010.

654 Odum, J. R., Jungkamp, T. P. W., Griffin, R. J., Forstner, H. J. L., Flagan, R. C., and Seinfeld, J. H.: Aromatics,
655 reformulated gasoline, and atmospheric organic aerosol formation, *Environmental Science & Technology*,
656 31, 1890-1897, 10.1021/es9605351, 1997.

657 Paulot, F., Crounse, J. D., Kjaergaard, H. G., Kurten, A., St Clair, J. M., Seinfeld, J. H., and Wennberg, P. O.:
658 Unexpected Epoxide Formation in the Gas-Phase Photooxidation of Isoprene, *Science*, 325, 730-733,
659 10.1126/science.1172910, 2009.

660 Paulsen, D., Dommen, J., Kalberer, M., Prevot, A. S. H., Richter, R., Sax, M., Steinbacher, M., Weingartner,
661 E., and Baltensperger, U.: Secondary organic aerosol formation by irradiation of 1,3,5-trimethylbenzene-
662 NO_x-H₂O in a new reaction chamber for atmospheric chemistry and physics, *Environmental Science &*
663 *Technology*, 39, 2668-2678, 10.1021/es0489137, 2005.

664 **Prenni, A. J., De Mott, P. J., and Kreidenweis, S. M.: Water uptake of internally mixed particles containing**
665 **ammonium sulfate and dicarboxylic acids, *Atmospheric Environment*, 37, 4243-4251, 10.1016/s1352-**
666 **2310(03)00559-4, 2003.**

667 Reinhardt, A., Emmenegger, C., Gerrits, B., Panse, C., Dommen, J., Baltensperger, U., Zenobi, R., and
668 Kalberer, M.: Ultrahigh mass resolution and accurate mass measurements as a tool to characterize oligomers
669 in secondary organic aerosols, *Analytical Chemistry*, 79, 4074-4082, 10.1021/ac062425v, 2007.

670 Rickard, A. R., Wyche, K. P., Metzger, A., Monks, P. S., Ellis, A. M., Dommen, J., Baltensperger, U., Jenkin,
671 M. E., and Pilling, M. J.: Gas phase precursors to anthropogenic secondary organic aerosol: Using the Master
672 Chemical Mechanism to probe detailed observations of 1,3,5-trimethylbenzene photo-oxidation,
673 *Atmospheric Environment*, 44, 5423-5433, 10.1016/j.atmosenv.2009.09.043, 2010.

674 Rincon, A. G., Guzman, M. I., Hoffmann, M. R., and Colussi, A. J.: Optical absorptivity versus molecular
675 composition of model organic aerosol matter, *Journal of Physical Chemistry A*, 113, 10512-10520,
676 10.1021/jp904644n, 2009.

677 Rodigast, M., Mutzel, A., Iinuma, Y., Haferkorn, S., and Herrmann, H.: Characterisation and optimisation of
678 a sample preparation method for the detection and quantification of atmospherically relevant carbonyl
679 compounds in aqueous medium, *Atmospheric Measurement Techniques*, 8, 2409-2416, 10.5194/amt-8-
680 2409-2015, 2015.

681 Sadezky, A., Chaimbault, P., Mellouki, A., Rompp, A., Winterhalter, R., Le Bras, G., and Moortgat, G. K.:
682 Formation of secondary organic aerosol and oligomers from the ozonolysis of enol ethers, *Atmospheric*
683 *Chemistry and Physics*, 6, 5009-5024, 2006.

684 Sareen, N., Schwier, A. N., Shapiro, E. L., Mitroo, D., and McNeill, V. F.: Secondary organic material formed
685 by methylglyoxal in aqueous aerosol mimics, *Atmospheric Chemistry and Physics*, 10, 997-1016, 2010.

686 Sato, K., Takami, A., Kato, Y., Seta, T., Fujitani, Y., Hikida, T., Shimono, A., and Imamura, T.: AMS and
687 LC/MS analyses of SOA from the photooxidation of benzene and 1,3,5-trimethylbenzene in the presence of
688 NO_x: effects of chemical structure on SOA aging, *Atmospheric Chemistry and Physics*, 12, 4667-4682,
689 10.5194/acp-12-4667-2012, 2012.

690 Saukko, E., Lambe, A. T., Massoli, P., Koop, T., Wright, J. P., Croasdale, D. R., Pedernera, D. A., Onasch, T.
691 B., Laaksonen, A., Davidovits, P., Worsnop, D. R., and Virtanen, A.: Humidity-dependent phase state of
692 SOA particles from biogenic and anthropogenic precursors, *Atmospheric Chemistry and Physics*, 12, 7517-
693 7529, 10.5194/acp-12-7517-2012, 2012.

694 Saxena, P., and Hildemann, L. M.: Water-soluble organics in atmospheric particles: A critical review of the
695 literature and application of thermodynamics to identify candidate compounds, *Journal of Atmospheric*
696 *Chemistry*, 24, 57-109, 10.1007/bf00053823, 1996.

697 Schaefer, T., van Pinxteren, D., and Herrmann, H.: Multiphase chemistry of glyoxal: Revised kinetics of the
698 alkyl radical reaction with molecular oxygen and the reaction of Glyoxal with OH, NO₃, and SO₄- in aqueous
699 solution, *Environmental Science & Technology*, 49, 343-350, 10.1021/es505860s, 2015.

700 Schwier, A. N., Sareen, N., Mitroo, D., Shapiro, E. L., and McNeill, V. F.: Glyoxal-Methylglyoxal Cross-
701 Reactions in Secondary Organic Aerosol Formation, *Environmental Science & Technology*, 44, 6174-6182,
702 10.1021/es101225q, 2010.

703 Sedehi, N., Takano, H., Blasic, V. A., Sullivan, K. A., and De Haan, D. O.: Temperature- and pH-dependent
704 aqueous-phase kinetics of the reactions of glyoxal and methylglyoxal with atmospheric amines and
705 ammonium sulfate, *Atmospheric Environment*, 77, 656-663, 10.1016/j.atmosenv.2013.05.070, 2013.

706 Seinfeld, J. H., Erdakos, G. B., Asher, W. E., and Pankow, J. F.: Modeling the formation of secondary organic
707 aerosol (SOA). 2. The predicted effects of relative humidity on aerosol formation in the α -pinene-, β -pinene-,
708 sabinene-, D3-carene-, and cyclohexene-ozone systems (vol 35, pg 1806, 2001), *Environmental Science &*
709 *Technology*, 35, 1806-1817, 10.1021/es011025g, 2001.

710 Smith, D. F., Kleindienst, T. E., and McIver, C. D.: Primary product distributions from the reaction of OH with
711 m-, p-xylene, 1,2,4- and 1,3,5-trimethylbenzene, *Journal of Atmospheric Chemistry*, 34, 339-364,
712 10.1023/A:1006277328628, 1999.

713 Song, C., Gyawali, M., Zaveri, R. A., Shilling, J. E., and Arnott, W. P.: Light absorption by secondary organic
714 aerosol from α -pinene: Effects of oxidants, seed aerosol acidity, and relative humidity, *Journal of*
715 *Geophysical Research-Atmospheres*, 118, 9, 10.1002/jgrd.50767, 2013.

716 Sun, Y. L., Zhang, Q., Anastasio, C., and Sun, J.: Insights into secondary organic aerosol formed via aqueous-
717 phase reactions of phenolic compounds based on high resolution mass spectrometry, *Atmospheric Chemistry*
718 *and Physics*, 10, 4809-4822, 10.5194/acp-10-4809-2010, 2010.

719 Surratt, J. D., Murphy, S. M., Kroll, J. H., Ng, N. L., Hildebrandt, L., Sorooshian, A., Szmigielski, R.,
720 Vermeylen, R., Maenhaut, W., Claeys, M., Flagan, R. C., and Seinfeld, J. H.: Chemical composition of
721 secondary organic aerosol formed from the photooxidation of isoprene, *Journal of Physical Chemistry A*,
722 110, 9665-9690, 10.1021/jp061734m, 2006.

723 Surratt, J. D., Kroll, J. H., Kleindienst, T. E., Edney, E. O., Claeys, M., Sorooshian, A., Ng, N. L., Offenberg,
724 J. H., Lewandowski, M., Jaoui, M., Flagan, R. C., and Seinfeld, J. H.: Evidence for organosulfates in
725 secondary organic aerosol, *Environmental Science & Technology*, 41, 517-527, 10.1021/es062081q, 2007.

726 Surratt, J. D., Chan, A. W. H., Eddingsaas, N. C., Chan, M. N., Loza, C. L., Kwan, A. J., Hersey, S. P., Flagan,
727 R. C., Wennberg, P. O., and Seinfeld, J. H.: Reactive intermediates revealed in secondary organic aerosol
728 formation from isoprene, *Proceedings of the National Academy of Sciences of the United States of America*,
729 107, 6640-6645, 10.1073/pnas.0911114107, 2010.

730 Szmigielski, R., Surratt, J. D., Vermeylen, R., Szmigielska, K., Kroll, J. H., Ng, N. L., Murphy, S. M.,
731 Sorooshian, A., Seinfeld, J. H., and Claeys, M.: Characterization of 2-methylglyceric acid oligomers in
732 secondary organic aerosol formed from the photooxidation of isoprene using trimethylsilylation and gas
733 chromatography/ion trap mass spectrometry, *Journal of Mass Spectrometry*, 42, 101-116, 10.1002/jms.1146,
734 2007.

735 Tan, Y., Carlton, A. G., Seitzinger, S. P., and Turpin, B. J.: SOA from methylglyoxal in clouds and wet
736 aerosols: Measurement and prediction of key products, *Atmospheric Environment*, 44, 5218-5226,
737 10.1016/j.atmosenv.2010.08.045, 2010.

738 Tan, Y., Lim, Y. B., Altieri, K. E., Seitzinger, S. P., and Turpin, B. J.: Mechanisms leading to oligomers and
739 SOA through aqueous photooxidation: insights from OH radical oxidation of acetic acid and methylglyoxal,
740 *Atmospheric Chemistry and Physics*, 12, 801-813, 10.5194/acp-12-801-2012, 2012.

741 Tilgner, A., and Herrmann, H.: Radical-driven carbonyl-to-acid conversion and acid degradation in
742 tropospheric aqueous systems studied by CAPRAM, *Atmospheric Environment*, 44, 5415-5422,
743 10.1016/j.atmosenv.2010.07.050, 2010.

744 Tolocka, M. P., Jang, M., Ginter, J. M., Cox, F. J., Kamens, R. M., and Johnston, M. V.: Formation of
745 oligomers in secondary organic aerosol, *Environmental Science & Technology*, 38, 1428-1434,
746 10.1021/es035030r, 2004.

747 Tuazon, E. C., MacLeod, H., Atkinson, R., and Carter, W. P. L.: α -dicarbonyl yields from the NO_x-air
748 photooxidations of a series of aromatic hydrocarbons in air, *Environmental Science & Technology*, 20, 383-
749 387, 10.1021/es00146a010, 1986.

750 **Virtanen, A., Joutsensaari, J., Koop, T., Kannosto, J., Yli-Pirila, P., Leskinen, J., Makela, J. M., Holopainen,**
751 **J. K., Poschl, U., Kulmala, M., Worsnop, D. R., and Laaksonen, A.: An amorphous solid state of biogenic**
752 **secondary organic aerosol particles, *Nature*, 467, 824-827, 10.1038/nature09455, 2010.**

753 Waxman, E. M., Elm, J., Kurten, T., Mikkelsen, K. V., Ziemann, P. J., and Vokamer, R.: Glyoxal and
754 Methylglyoxal Setschenow Salting Constants in Sulfate, Nitrate, and Chloride Solutions: Measurements and
755 Gibbs Energies, *Environmental Science & Technology*, 49, 11500-11508, 10.1021/acs.est.5b02782, 2015.

756 Wyche, K. P., Monks, P. S., Ellis, A. M., Cordell, R. L., Parker, A. E., Whyte, C., Metzger, A., Dommen, J.,
757 Duplissy, J., Prevot, A. S. H., Baltensperger, U., Rickard, A. R., and Wulfert, F.: Gas phase precursors to
758 anthropogenic secondary organic aerosol: detailed observations of 1,3,5-trimethylbenzene photooxidation,
759 *Atmospheric Chemistry and Physics*, 9, 635-665, 10.5194/acp-9-635/2009/, 2009.

760 Yasmeen, F., Sauret, N., Gal, J. F., Maria, P. C., Massi, L., Maenhaut, W., and Claeys, M.: Characterization
761 of oligomers from methylglyoxal under dark conditions: a pathway to produce secondary organic aerosol

762 through cloud processing during nighttime, *Atmospheric Chemistry and Physics*, 10, 3803-3812,
763 10.5194/acp-10-3803-2010, 2010.

764 Yu, J. Z., Jeffries, H. E., and Sexton, K. G.: Atmospheric photooxidation of alkylbenzenes -I. Carbonyl product
765 analyses, *Atmospheric Environment*, 31, 2261-2280, 10.1016/S1352-2310(97)00011-3, 1997.

766 Zahardis, J., LaFranchi, B. W., and Petrucci, G. A.: Photoelectron resonance capture ionization-aerosol mass
767 spectrometry of the ozonolysis products of oleic acid particles: Direct measure of higher molecular weight
768 oxygenates, *Journal of Geophysical Research-Atmospheres*, 110, 10, 10.1029/2004jd005336, 2005.

769 Zappoli, S., Andracchio, A., Fuzzi, S., Facchini, M. C., Gelencser, A., Kiss, G., Krivacsy, Z., Molnar, A.,
770 Meszaros, E., Hansson, H. C., Rosman, K., and Zebuhr, Y.: Inorganic, organic and macromolecular
771 components of fine aerosol in different areas of Europe in relation to their water solubility, *Atmospheric
772 Environment*, 33, 2733-2743, 10.1016/s1352-2310(98)00362-8, 1999.

773 Zhao, J., Levitt, N. P., Zhang, R. Y., and Chen, J. M.: Heterogeneous reactions of methylglyoxal in acidic
774 media: Implications for secondary organic aerosol formation, *Environmental Science & Technology*, 40,
775 7682-7687, 10.1021/es060610k, 2006.

776 Ziemann, P. J.: Atmospheric chemistry Phase matters for aerosols, *Nature*, 467, 797-798, 2010.

777 Zuend, A., Marcolli, C., Peter, T., and Seinfeld, J. H.: Computation of liquid-liquid equilibria and phase
778 stabilities: implications for RH-dependent gas/particle partitioning of organic-inorganic aerosols,
779 *Atmospheric Chemistry and Physics*, 10, 7795-7820, 10.5194/acp-10-7795-2010, 2010.

Table 1: Overview of methods for structure elucidation of oligomeric compounds (after Hallquist et al., 2009).

Method	Reference
(Matrix assisted) laser desorption/ionisation mass spectrometry (MALDI-MS, LDI-MS)	Dommen et al., 2006; Kalberer et al., 2004; Kalberer et al., 2006; Reinhardt et al., 2007; Holmes and Petrucci, 2006; Surratt et al., 2006; Denkenberger et al., 2007
Fourier transform ion cyclotron resonance mass spectrometry (FT-ICR-MS)	Kundu et al., 2012; Altieri et al., 2008; Tolocka et al., 2004; Hall and Johnston, 2012; Denkenberger et al., 2007; Tan et al., 2012
On-line atmospheric pressure chemical ionisation tandem mass spectrometry (APCI tandem MS)	Müller et al., 2008
Aerosol mass spectrometry (AMS)	Sareen et al., 2010; Schwier et al., 2010; Bahreini et al., 2005; Heaton et al., 2007
Electrospray ionisation mass spectrometry (ESI/MS, ESI/MS/MS)	Altieri et al., 2008; Hall and Johnston, 2012; Surratt et al., 2006; Yasmeen et al., 2010; Hastings et al., 2005; Bones et al., 2010; Surratt et al., 2007; Hamilton et al., 2006; Sadezky et al., 2006; Sato et al., 2012; Noziere et al., 2010; Tolocka et al., 2004; Iinuma et al., 2004; Nguyen et al., 2011; Bahreini et al., 2005
Aerosol time of flight mass spectrometry (ALTOFMS)	Huang et al., 2015
Gas chromatography mass spectrometry (GC/MS)	Hastings et al., 2005; Surratt et al., 2006; Szmigielski et al., 2007; Angove et al., 2006
Ion trap mass spectrometry (IT-MS)	Surratt et al., 2006; Gao et al., 2004
Photoelectron resonance capture ionisation-aerosol mass spectrometry (PERCI-MS)	Zahardis et al., 2005
Ultraviolet-visible spectroscopy (UV/Vis)	Nemet et al., 2004; Noziere and Esteve, 2005; Bones et al., 2010; Song et al., 2013; Casale et al., 2007; Alfarra et al., 2006; Drozd and McNeill, 2014; Noziere and Cordova, 2008
Fourier transform infrared spectroscopy (FTIR)	Loeffler et al., 2006; Bones et al., 2010; Jang et al., 2003; Jang and Kamens, 2001; Holmes and Petrucci, 2006
nuclear magnetic resonance spectroscopy (NMR)	Nemet et al., 2004; Bones et al., 2010; Angove et al., 2006; Garland et al., 2006; Kua et al., 2013; De Haan et al., 2011

Table 2: Experiments in the aerosol chamber LEAK for the OH radical oxidation of 1,3,5-TMB. All experiments were conducted at 293 K and with 91.9 ppb 1,3,5-TMB.

Experiment number	seed	RH [%]	O _{3ini} [ppb]	ΔHC [ppb]	ΔM [μg m ⁻³]	SOA yield Y _{SOA} [%]
#1	78 mmol L ⁻¹ NH ₄ H ₂ SO ₄	≈ 0	≈ 137	57.2	19.7	7.00
#2	78 mmol L ⁻¹ NH ₄ H ₂ SO ₄	50	≈ 133	55.2	11.3	4.12
#3	78 mmol L ⁻¹ NH ₄ H ₂ SO ₄	75	≈ 134	56.2	14.2	5.14
#4	60 mmol L ⁻¹ (NH ₄) ₂ SO ₄ / 0.4 mmol L ⁻¹ H ₂ SO ₄	≈ 0	≈ 132	56.2	18.1	6.55
#5	60 mmol L ⁻¹ (NH ₄) ₂ SO ₄ / 0.4 mmol L ⁻¹ H ₂ SO ₄	50	≈ 135	56.5	11.7	4.21
#6	60 mmol L ⁻¹ (NH ₄) ₂ SO ₄ / 0.4 mmol L ⁻¹ H ₂ SO ₄	75	≈ 144	57.2	13.9	4.94

1,3,5-TMB: 1,3,5-Trimethylbenzene; RH: Relative humidity

Table 3: Investigated method parameters and used PTFE filters from 1,3,5-TMB oxidation for method development.

Parameter	Range	Filter of experiment
Heating time	15, 24 , 30, 48 hours	#3
pH	1 , 3, 5, 7	#2
Heating temperature	50, 100 °C	#2

Selected parameters given in **bold**.

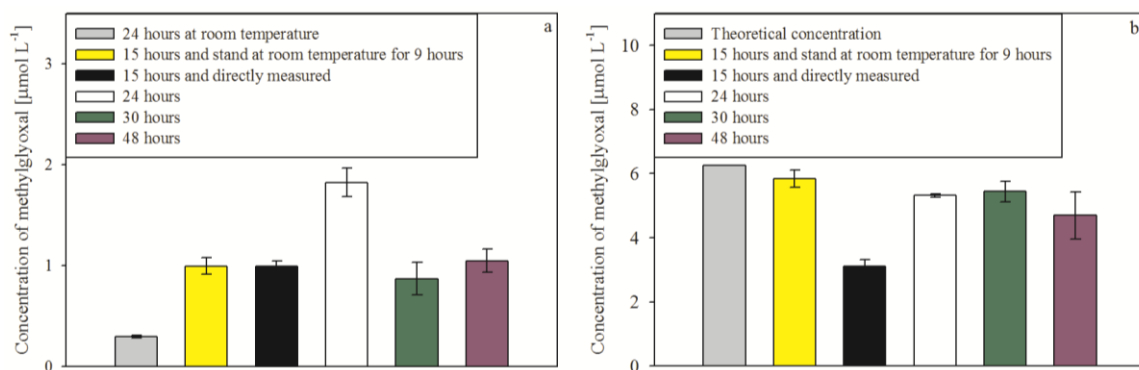


Figure 1: Influence of the heating time on the detected methylglyoxal concentrations in filter samples (a) and the methylglyoxal standard solution (b).

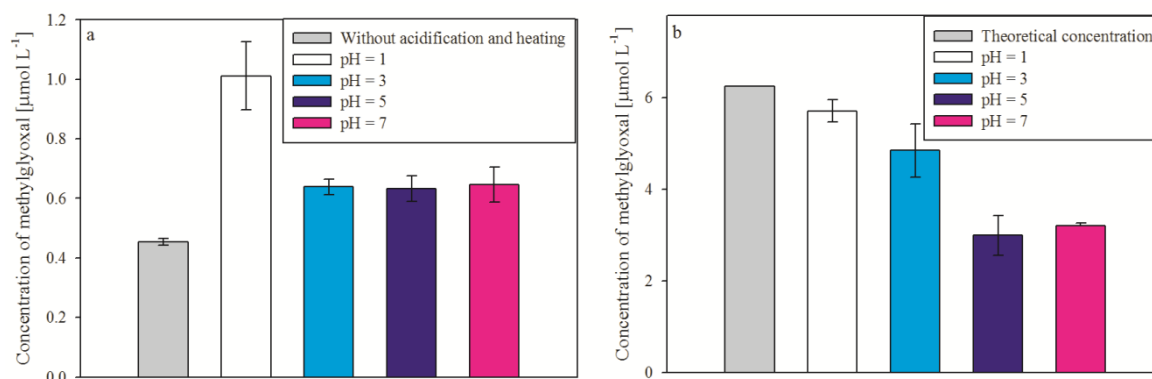


Figure 2: Influence of pH on the detected methylglyoxal concentration in filter samples (a) and the methylglyoxal standard solution (b).

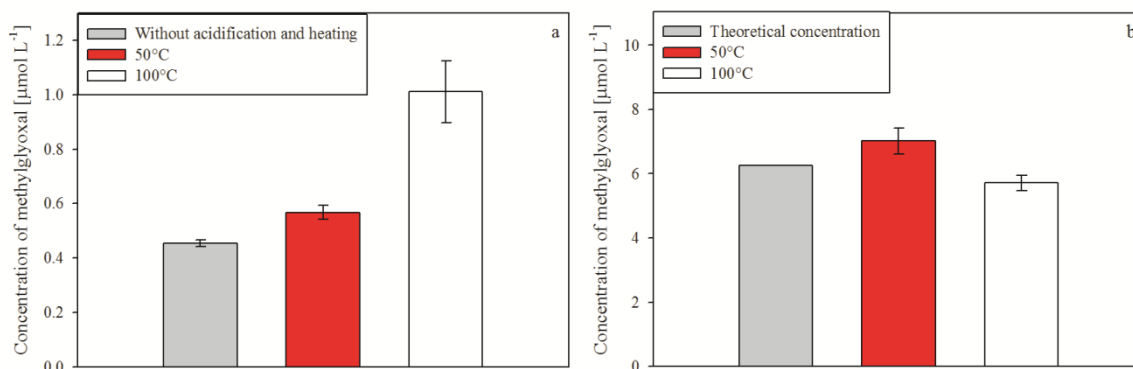


Figure 3: Influence of the heating temperature on the detected methylglyoxal concentration in filter samples (a) and the standard solution (b).

Table 4: SOA yields (Y_{SOA}) of 1,3,5-TMB reported in the literature.

SOA yield Y_{SOA} [%]	Reference
2.81 - 7.91	Cocker et al., 2001
4.5 - 8.34	Healy et al., 2008
$2.5 \pm 0.1 - 15.6 \pm 1.0$	Sato et al., 2012
3.1	Odum et al., 1997
0.41 ± 0.1	Kleindienst et al., 1999
0.29 - 6.36	Wyche et al., 2009
4.7 ± 0.7	Paulsen et al., 2005
$7.1 \pm 0.3 - 13.8 \pm 0.6$	Cao and Jang, 2007
4.12 - 7.00	This work

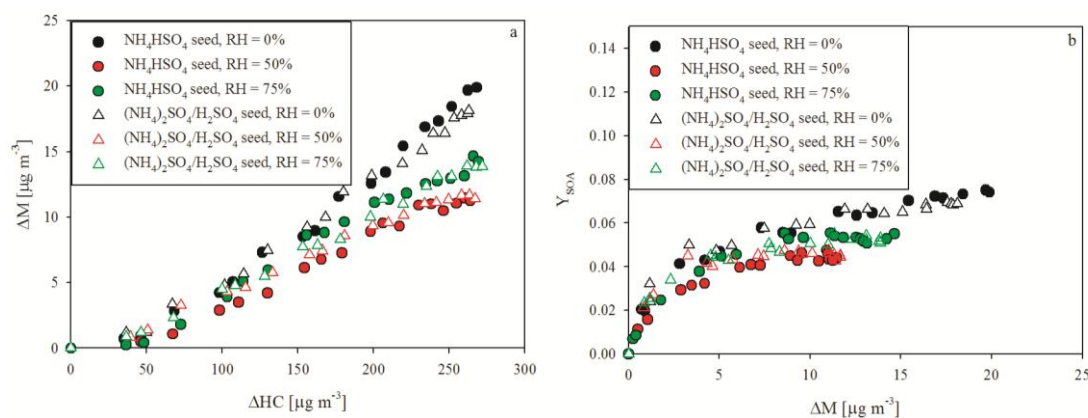


Figure 4: SOA growth curve (a) and yield curve (b) of the OH radical oxidation of 1,3,5-TMB in the presence of NH_4HSO_4 or $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed under variation of RH.

Table 5: LWC and pH of the seed particles calculated with E-AIM.

seed	RH [%]	pH _{seed} ^[a]	LWC [g m ⁻³] ^[a]
78 mmol L ⁻¹ NH_4HSO_4	≈ 0	— ^[b]	— ^[b]
78 mmol L ⁻¹ NH_4HSO_4	50	0.1	6.66×10^{-6}
78 mmol L ⁻¹ NH_4HSO_4	75	1.2	12.29×10^{-6}
60 mmol L ⁻¹ $(\text{NH}_4)_2\text{SO}_4/0.4$ mmol L ⁻¹ H_2SO_4	≈ 0	— ^[b]	— ^[b]
60 mmol L ⁻¹ $(\text{NH}_4)_2\text{SO}_4/0.4$ mmol L ⁻¹ H_2SO_4	50	4.0	4.25×10^{-6}
60 mmol L ⁻¹ $(\text{NH}_4)_2\text{SO}_4/0.4$ mmol L ⁻¹ H_2SO_4	75	4.2	11.56×10^{-6}

LWC: Liquid water content; [a] pH and LWC of the seed particles were calculated for different RH using model II from the extended aerosol thermodynamic model (E-AIM; Clegg et al., 1998); [b] calculation of the pH and LWC was not possible due to the low relative humidity of $\text{RH} \approx 0\%$. A $\text{RH} = 10\%$ is set as lower limit in E-AIM.

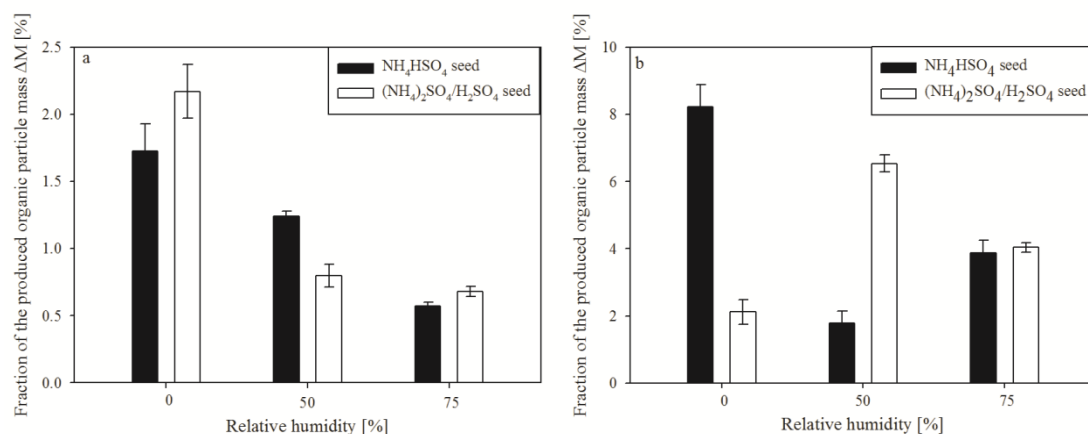


Figure 5: Contribution of methylglyoxal (a) and their heat-decomposable oligomers (b) to the produced organic particle mass (ΔM) with NH_4HSO_4 and $(\text{NH}_4)_2\text{SO}_4/\text{H}_2\text{SO}_4$ seed particles under variation of the relative humidity.