

Responses to Referee #1

We thank Referee #1 for the helpful comments. Please find our responses below.

Major issues

1. I find it confusing that the manuscript switches between cloud droplets and aerosol particles. It looks like measurements of aerosols are compared to model results of cloud chemistry. Is this the case? If yes, please explain why this approach can be valid.

Response: This is not the case. Sulfate aerosol in the environment is produced via gas-phase and aqueous-phase chemistry, the latter occurring primarily in clouds. The model is used to distinguish between the different sulfate aerosol production mechanisms (gas, aqueous, heterogeneous) in order to interpret the observations of sulfate isotopes. We have incorporated the following changes to the manuscript to help make this more clear.

Changes to the manuscript: Add “The model calculates $\Delta^{17}\text{O}$ of bulk sulfate in the model grid box ($\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$) based on the relative importance of each sulfate production mechanism for total sulfate abundance. This is then compared to the $\Delta^{17}\text{O}$ measurement of bulk sulfate collected on aerosol filters ($\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$) to investigate sulfate formation mechanisms in the MBL (Sect. 4.2).” to Line 180, in Section 3.

2. In Table 1, you cite Fogelman et al. (1989) and Troy and Margerum (1991) as references for $\Delta^{17}\text{O} = 0$ for the reactions $\text{S(IV)} + \text{HOX}$. However, I do not see any isotope chemistry in these papers. How do you derive this zero value? How would your results change if $\Delta^{17}\text{O}$ is non-zero for $\text{S(IV)} + \text{HOX}$?

Response: The Eqs. (2-3) in Fogelman et al. (1989) and Eqs. (5-6) in Troy and Margerum (1991) show that reactions between HOX and SO_3^{2-} occur via nucleophilic attack of SO_3^{2-} onto the X atom of HOX (X = Br or Cl), followed by rapid hydrolysis of XSO_3^- . This is shown in Eqs. (1-2) in our paper. The mechanism can be used to determine the resulting $\Delta^{17}\text{O}$ value of sulfate. We can see that the three oxygen atoms of XSO_3^- are from SO_3^{2-} , which will have a $\Delta^{17}\text{O}$ value of 0 ‰ via isotopic exchange of S(IV) with water. Hydrolysis adds one oxygen atom from H_2O to XSO_3^- to form SO_4^{2-} . Since all four oxygen atoms of sulfate are derived from water, the resulting

$\Delta^{17}\text{O}$ value of sulfate will be the same as water, which is 0 ‰. As long as SO_3^{2-} does not get an oxygen atom directly from HOX to form SO_4^{2-} , the $\Delta^{17}\text{O}$ of SO_4^{2-} formed via reaction “ $\text{SO}_3^{2-} + \text{HOX}$ ” will be 0 ‰. Liu (2000) suggests the reaction of HOBr with HSO_3^- follows a similar pathway as with SO_3^{2-} , i.e. nucleophilic attack of SO_3^{2-} onto the Br atom of HOBr, followed by rapid hydrolysis of BrSO_3^- . Thus, the $\Delta^{17}\text{O}$ of SO_4^{2-} formed via reaction “ $\text{HSO}_3^- + \text{HOBr}$ ” will also be 0 ‰. There is no laboratory study on the reaction “ $\text{HSO}_3^- + \text{HOCl}$ ”, but it is reasonable to assume the same reaction mechanism as for these other pathways leading to $\Delta^{17}\text{O} = 0$ ‰.

For $\Delta^{17}\text{O}_{\text{HOX}}(\text{SO}_4^{2-})$ to be non-zero, the mechanism for sulfate formation via HOX oxidation would need to follow direct transfer of the oxygen atom of HOX to S(IV). This would lead to $\Delta^{17}\text{O}_{\text{HOX}}(\text{SO}_4^{2-})$ values similar to that from O_3 oxidation, and thus HOX oxidation would not resolve the model-measurement discrepancy. However, the above-mentioned laboratory studies all suggest that this is not the case.

We have made the following changes to the manuscript to make it more clear how we determine $\Delta^{17}\text{O}_{\text{HOX}}(\text{SO}_4^{2-})$:

Changes to the manuscript: At Line 104, change “Since HOX promotes sulfate formation by adding one oxygen atom from H_2O to sulfate instead of transferring its own oxygen atom (Fogelman et al., 1989; Troy and Margerum, 1991; Yiin and Margerum, 1988), the $\Delta^{17}\text{O}$ of sulfate produced from HOX oxidation is expected to be 0 ‰.” to “Since HOX promotes sulfate formation via “ $\text{SO}_3^{2-} + \text{HOX}$ ” reactions by adding one oxygen atom from H_2O to sulfate instead of transferring its own oxygen atom (Fogelman et al., 1989; Troy and Margerum, 1991; Yiin and Margerum, 1988), the $\Delta^{17}\text{O}$ of sulfate produced from “ $\text{SO}_3^{2-} + \text{HOX}$ ” reactions is expected to have the same $\Delta^{17}\text{O}$ as water (0 ‰) (Gonfiantini et al., 1993). Liu (2000) suggests the reaction of HOBr with HSO_3^- follows a similar pathway as with SO_3^{2-} (Eqs. 1-2), resulting in $\Delta^{17}\text{O}$ of 0 ‰ for sulfate produced via this reaction. We assume that the reaction of HOCl with HSO_3^- follows a similar pathway as the reaction of HOBr with HSO_3^- (Eqs. 3-4) and produces sulfate with $\Delta^{17}\text{O}$ of 0 ‰. Based on these mechanistic studies, the $\Delta^{17}\text{O}$ of sulfate produced from HOX oxidation is expected to be 0 ‰.”.

3. You conclude that a fraction of 33 to 50 % of the sulfate can be produced by about 0.1 pmol/mol HOX. Vogt et al. (1996), however, needed 100 times more HOX (10 pmol/mol) to achieve a similar fraction (60 %). Why is there such a big difference

between these two studies? Can the difference be explained by different rate coefficients that were used? Were different reactions included in these studies? If not, what else could it be?

Response: The rate constants used in our calculations are $k_{\text{HOX}+\text{SO}_3^{2-}} = 2.9 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ and $k_{\text{HOX}+\text{HSO}_3^-} = 2.0 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$. In comparison, the rate constants used in Vogt et al. (1996) are $k_{\text{HOCl}+\text{SO}_3^{2-}} = k_{\text{HOCl}+\text{HSO}_3^-} = 7.6 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$ and $k_{\text{HOBr}+\text{SO}_3^{2-}} = k_{\text{HOBr}+\text{HSO}_3^-} = 5.0 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$. Thus, we have used higher rate constants for “S(IV)+HOCl” and lower rate constants for “S(IV)+HOBr”. Given the similarity in the overall rate constants, the HOX concentration difference in Vogt et al. (1996) and our study is unlikely to be explained by the difference in rate constants.

The difference of calculated HOX concentration between our study and Vogt et al. (1996) is caused by several reasons described below:

(1) We use a higher Henry’s law constant for HOX ($H_{\text{HOX}}=975 \text{ M atm}^{-1}$), which will require a lower gas-phase HOX concentration to produce the same amount of sulfate. In Vogt et al. (1996), the Henry’s law constant for HOCl (H_{HOCl}) and HOBr (H_{HOBr}) is 660 M atm^{-1} and 93 M atm^{-1} , respectively. The updated Henry’s law constants used in our study requires about an order of magnitude lower concentrations of gas-phase HOX.

(2) The averaged H_2O_2 and O_3 concentrations ($[\text{H}_2\text{O}_2]=0.6 \text{ ppt}$, $[\text{O}_3]=18 \text{ ppb}$) in our studies are lower. In Vogt et al. (1996), $[\text{H}_2\text{O}_2]$ and $[\text{O}_3]$ are 0.8 ppt and 40 ppb , respectively. Thus, compared to our study, they need higher HOX concentrations to compete with H_2O_2 and O_3 to produce sulfate.

(3) The HOX concentration of 0.1 ppt in our study is calculated by assuming equilibrium between gaseous and aqueous phase of HOX. The calculation of HOX in Vogt et al. (1996) considers all mass transfer limitations. It is possible that aqueous-phase HOX concentrations are lower than their equilibrium concentrations due to fast chemical removal in the aqueous-phase. In Vogt et al. (1996), HOBr is removed by both Cl^- and S(IV) in the aqueous phase. If Cl^- concentrations are high enough, it will compete with “S(IV) + HOBr” for loss of HOBr. Von Glasow et al. (2002) has shown that HOBr becomes more important than HOCl in terms of S(IV) oxidation in clouds where LWC is high and Cl^- is diluted.

(4) The HOX concentration of 0.1 ppt calculated in our study is a 24-hour-mean concentration. The HOX concentration in Vogt et al. (1996) is about 10 ppt during the day, but is 2 orders of magnitude lower at night. Thus, the daily mean HOX concentration is on the order of several ppt in Vogt et al. (1996), rather than 10 ppt stated in the comment above.

We have added the following discussion to the manuscript:

Changes to the manuscript: Add one paragraph to Line 355: “Uncertainties in our calculated $[\text{HOX}]_g$ can originate from (1) uncertainties in the rate constant for reaction between HOX and HSO_3^- ($k_{\text{HOX}+\text{HSO}_3^-}$), (2) uncertainties in the Henry’s law constant for HOX (H_{HOX}) and (3) the efficiency of reactive uptake of gas-phase HOX onto cloud droplets that is not accounted for in our assumption of equilibrium of HOX between the gas and aqueous phase.

In comparison, a box-modeling study by Vogt et al. (1996) estimated that daytime-averaged $[\text{HOX}]_g$ on the order of 10 ppt is needed to achieve a similar fraction ($\approx 60\%$) of sulfate formed via HOX oxidation. The difference in $[\text{HOX}]_g$ between our study and Vogt et al. (1996) is caused by several factors. First, H_{HOBr} in our calculations is an order of magnitude larger than that in Vogt et al. (1996), so that our calculations require an order of magnitude lower $[\text{HOBr}]_g$ to produce the same aqueous-phase concentration. Second, H_2O_2 and O_3 mixing ratios in our calculations (Eq. 9) are lower than those in Vogt et al. (1996) (0.6 ppt vs. 0.8 ppt for $[\text{H}_2\text{O}_2]$ and 18 ppt vs. 40 ppt for $[\text{O}_3]$), so that higher $[\text{HOBr}]_g$ is needed in Vogt et al. (1996) to compete with S(IV) oxidation by H_2O_2 and O_3 . Third, in our simple calculation we assume equilibrium of HOX between the gas- and aqueous-phase while Vogt et al. (1996) considers all mass transfer limitations. Higher $[\text{HOX}]_g$ will be calculated if diffusion and subsequent mass accommodation of gas-phase HOX onto the cloud droplets is not fast enough to compensate for the loss of HOX from aqueous-phase chemistry. Fourth, $[\text{HOX}]_g$ on the order of 0.1 ppt calculated in our study is a daily-averaged concentration while $[\text{HOX}]_g$ on the order of 10 ppt in Vogt et al. (1996) is a daytime-averaged concentration, and hence are not directly comparable. Vogt et al. (1996) has shown that nighttime-averaged $[\text{HOX}]_g$ is about two orders of magnitude lower than daytime-averaged $[\text{HOX}]_g$.”

Delete “In comparison,” at Line 356.

Specific Comments

1. Abstract: I think it should be mentioned here that uncertainties in reactive halogen concentrations are probably the main reason why halogen chemistry is excluded in large scale models (apart from uncertainties in the reaction rates).

Response: Thanks for the comment. We agree that the largest uncertainty in reaction rates is due to the uncertainty in the concentrations of HOX. We have included the following sentence in our abstract.

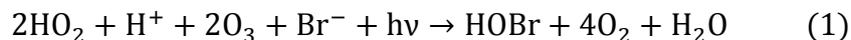
Changes to the manuscript: Change “a production mechanism typically excluded in large scale models due to uncertainties in the reaction rates” to “a production mechanism typically excluded in large scale models due to uncertainties in the reaction rates, which are due mainly to uncertainties in reactive halogen concentrations” at Line 23.

2. page 5, line 131: Here you list several ions which were measured. Unfortunately, the list does not include bromide. Since the reaction $\text{HOBr} + \text{S(IV)}$ produces bromide, it would be very interesting to know bromide aerosol concentrations. Would it be possible to analyze the samples for bromide?

Response: We have measured Br^- concentration ($[\text{Br}^-]$) (unit: $\mu\text{g}/\text{m}^3$) for the Xue-Long cruise samples but not for the Malaspina cruise samples. Fig. 1 shows the relationship between our observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ and $[\text{Br}^-]$ concentrations, and the relationship between our calculated $[\text{HOX}]_{\text{g}}$ concentrations and observed $[\text{Br}^-]$ concentrations. There is no apparent relationship between $\Delta^{17}\text{O}(\text{SO}_4^{2-})$ and $[\text{Br}^-]$ or between $[\text{HOX}]_{\text{g}}$ and $[\text{Br}^-]$. One would not necessarily expect a relationship between $[\text{Br}^-]$ and $[\text{HOX}]_{\text{g}}$ because the production of reactive halogens is dependent on additional factors other than Br^- , as discussed below.

We have investigated this relationship further using model results from Schmidt et al. (2016). Fig. 2 shows the global distribution of 1st-model-level (<100m) HOBr abundance and Br^- abundance in January 2007 from GEOS-Chem model results in Schmidt et al. (2016). A scatter plot showing the relationship between HOBr and Br^- is shown in Fig. 3. A significant correlation between HOBr and Br^- in the model is not found. This is because the amount of HOBr produced from Br^- and thus the amount of sulfate produced from HOBr depends on other factors in addition to Br^- concentration. These factors include (1) the acidity of aerosols because regeneration of HOBr is acid catalyzed (Fickert et al., 1999), (2) the amount of ultraviolet radiation

that photolyzes HOBr, (3) the concentrations of HO₂ and O₃ that produce HOBr (Eq. 1) and (4) the concentrations of species such as Cl⁻ and Br⁻ that act as a sink for HOBr. Thus, [Br⁻] alone is not a good proxy for the “HOBr + S(IV)” reaction.



Changes to the manuscript: Add “Bromide aerosol concentrations ([Br⁻]) were also measured for the Xue-Long samples (Supplementary data). There is no relationship between observed Br⁻ concentration and Δ¹⁷O(SO₄²⁻) nor our calculated [HOX]_g (not shown) because factors such as aerosol pH, sunlight and oxidants play an important role in the acid-catalyzed formation of reactive halogens and removal of HOBr (Fickert et al., 1999; Schmidt et al., 2016), Similarly, there is no relationship between [Br⁻] and HOBr mixing ratios in the global modeling study by Schmidt et al. (2016) (not shown). Thus, [Br⁻] alone is not a good proxy for the “HOBr + S(IV)” reaction.” to Line 135 in Sect. 2.

We have also added the observed [Br⁻] to the Supplementary data.

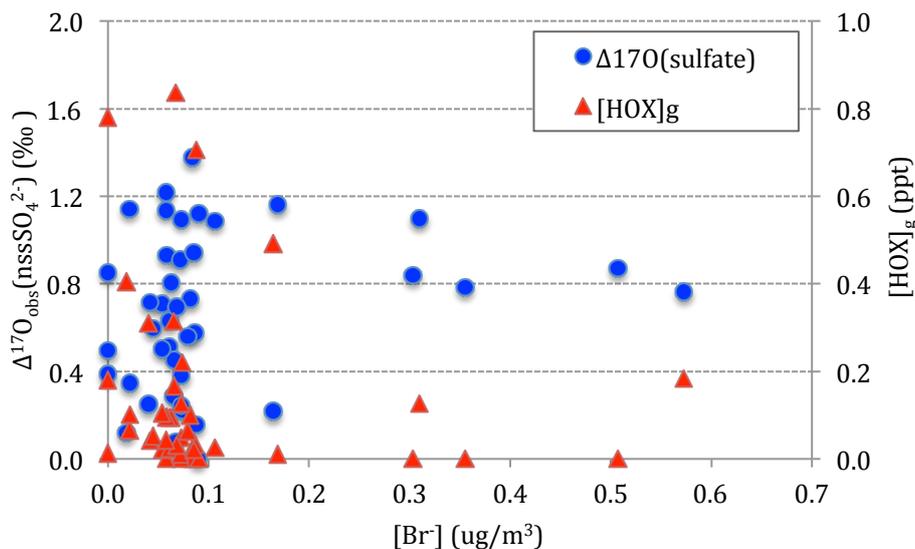


Fig. 1 The Δ¹⁷O(nssSO₄²⁻), [HOX]_g and [Br⁻] for each sample from the Xue-Long cruise.

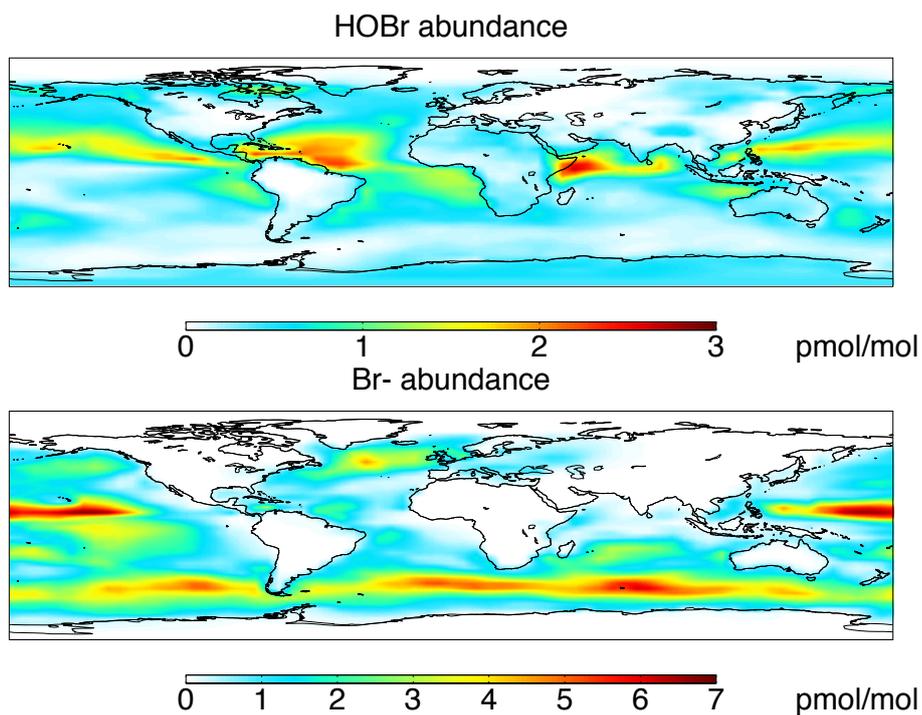


Fig. 2 Global distribution of 1st-model-level (<100m) HOBr abundance and Br⁻ abundance in January 2007 from GEOS-Chem model outputs in Schmidt et al. (2016).

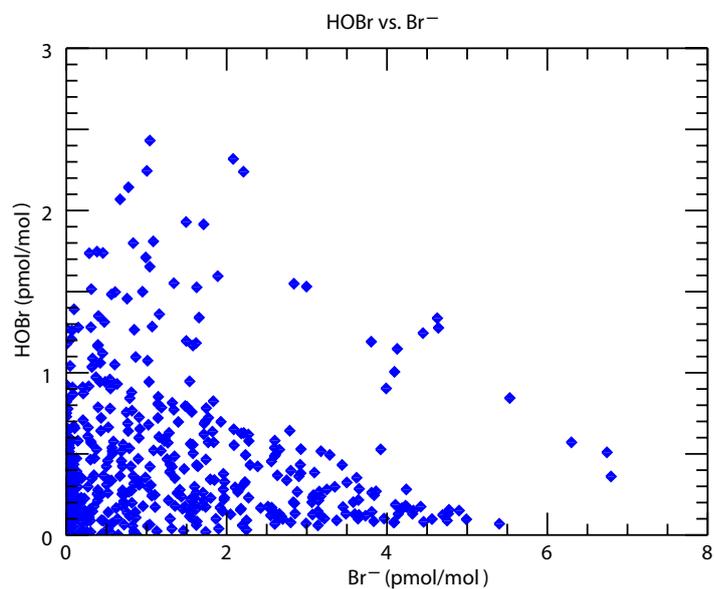


Fig. 3 The relationship between HOBr and Br⁻ in Fig. 2.

3. page 6, line 173: Metal-catalyzed oxidation of S(IV) is mentioned here. Could you add the information which metals (Fe?, Mn?) are included in the model and how the model calculates their concentrations? I think this information is necessary to understand why this pathway contributes less than 1 % to S(IV) oxidation in the model.

Response: Our model includes Fe and Mn for S(IV) oxidation. We have added this information to the manuscript accordingly, which is presented in more detail in Alexander et al. (2009). The reason why metal-catalyzed oxidation pathway is not important in our samples is that our samples are collected in the remote marine boundary layer where trace metals concentrations are low due to the distance from both anthropogenic (coal combustion) and natural (dust) sources of these metals.

Changes to the manuscript: Add “The parameterization of the metal-catalyzed S(IV) oxidation pathway is described in detail in Alexander et al. (2009). The trace metals included are Fe and Mn, whose oxidation states Fe(III) and Mn(II) catalyze S(IV) oxidation. Soil-derived Fe ($[Fe]_{soil}$) is scaled to modeled dust concentration as 3.5% of total dust mass while soil-derived Mn ($[Mn]_{soil}$) is a factor of 50 lower than $[Fe]_{soil}$. Anthropogenic Mn ($[Mn]_{anthro}$) is scaled as 1/300 of primary sulfate concentration while anthropogenic Fe ($[Fe]_{anthro}$) is 10 times that of $[Mn]_{anthro}$. We assume that 50% of Mn is dissolved in cloud water as Mn(II) oxidation state. For Fe, we assume that 10% of $[Fe]_{anthro}$ and 1% of $[Fe]_{soil}$ is dissolved in cloud water. 10% of the dissolved Fe is in Fe(III) oxidation state during daytime and 90% at night.” to Line 173.

Technical Comments

1. The physical properties “mixing ratio” and “concentration” are used as if they were identical. This is not the case! (for details, see <http://www.rolf-sander.net/res/vol1kg.pdf>) Please check all occurrences of the word “concentration” in the main text and check if it should read “mixing ratio” instead.

Response: Thanks for pointing this out. We have updated our manuscript accordingly.

Changes to the manuscript: Change “concentration” into “mixing ratio” at Lines 24, 82-89, 338, 351, 354-369, 379, 391-395 and 447-448.

2. page 13, line 366: If Le Breton et al. is still in review, it should be called (2016) not (2015).

Response: Thanks for pointing this out. We have updated our manuscript accordingly.

Changes to the manuscript: Change “2015” into “2016” at Lines 88-89, 366 and 563.

3. page 16, line 439: The unit “permil” is missing for $\Delta^{17}\text{O}$.

Response: Thanks for pointing this out. We have updated our manuscript accordingly.

Changes to the manuscript: Add “‰” to $\Delta^{17}\text{O}$ at Line 439.

Responses to Referee #2

We thank Referee #2 for the helpful comments. Please find our responses below.

1. First, in paragraph 230, the fractional contributions are discussed rather than the concentration. Later in the manuscript, the authors address the concentration but it may be worth mentioning the rationale for describing fractional contribution rather than amount here.

Response: Thanks for this suggestion. We separated the “Result” section (Sect. 4) and “Discussion” section (Sect. 5) in the manuscript. For the “Result” section, we showed only the observations of concentrations and isotopes and primary modeling results (without HOX contribution). Based on these results, we calculated the fractional contribution of sulfate formation from HOX and $[\text{HOX}]_g$, which is now described in the “Discussion” section. In paragraph 230, all the fractional contributions shown are obtained directly from the GEOS-Chem model, which does not include the fractional contribution from HOX and $[\text{HOX}]_g$ mixing ratio.

Additional changes to the manuscript: Add “Based on the modeled fractional contributions to the sulfate burden from each oxidant except HOX, and knowledge about reaction rate constants of “HOX + S(IV)” reactions and Henry’s law constants of HOX, we calculate the amount of HOX needed to explain the discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ in Sect. 5.2.4.” to Line 234 in Section 4.2.

2. Second, dark OH reactions (from nitrogen reactions on aerosol surfaces, e.g. Fuchs et al., 2013 doi:10.1038/ngeo1964) has recently been described as an important nighttime oxidation pathway that is typically not considered in chemistry and aerosol models. What would the implications be here and can it be ignored?

Response: Thanks for raising this interesting question. High nighttime OH concentrations have been observed in forests (Faloona et al., 2001) and urban areas (Lu et al., 2014). The mechanism behind this is still unknown. Possible explanations include unimolecular reactions of isoprene-derived peroxy radicals (Fuchs et al., 2013) and an additional RO_x production process from VOCs and additional recycling ($\text{RO}_2 \rightarrow \text{HO}_2 \rightarrow \text{OH}$) (Lu et al., 2014).

In the offline aerosol version of GEOS-Chem that we used in this study, nighttime OH concentration are set equal to zero, so that nighttime oxidation of S(IV) by OH does not occur in the model. However, as the gas-phase reaction “SO₂ + OH” is relatively slow compared to aqueous S(IV) oxidation, adding this dark OH production mechanism will be negligible. In our sensitivity study of doubling OH concentrations shown in Sect. 5.2.1, the fraction of sulfate produced by OH oxidation increases only from 20% to 27% globally. In addition, we think this dark OH production can be ignored in our study because most of our samples are from remote marine boundary layer where isoprene and VOCs abundance are low.

Changes to the manuscript: Add “A diurnal variation as a function of solar zenith angle is applied to OH concentrations and photolytic loss rates of H₂O₂ in the model. NO₃ is set to be zero during daytime.” to Line 172 in Sect. 3.

Add “Thus, doubling OH concentrations has an insignificant effect on $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$. The nighttime OH concentrations observed in forests and urban areas (Faloona et al., 2001; Lu et al., 2014) should have insignificant effects on our model results as most of our samples are collected in the remote MBL where isoprene and VOCs abundance are low. Our sensitivity study with doubled OH suggests additional nighttime sources of OH are not able to resolve the modeled overestimate of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ observations.” to Line 277 in Sect. 5.2.1.

3. Third, how realistic is it that S(IV) + HOX results in no pH change in clouds (paragraph 420) and although this is the treatment used in the model, what are the difficulties associated with changing both cloud pH and the fraction of S(IV) + HOX at the same time?

Response: We do not have “S(IV) + HOX” reaction in the model. Adding “S(IV) + HOX” reaction in the model will probably decrease cloud pH, which will result in decreases in the fractional contribution of O₃ relative to H₂O₂ ($f_{\text{O}_3}/f_{\text{H}_2\text{O}_2}$) (Line 422). Thus, the HOBr concentration calculated by assuming no change in pH is likely an overestimate (Line 432). We do not have enough information to estimate the magnitude of pH change after adding “S(IV) + HOX” reaction.

For a follow-up study, we are implementing “S(IV) + HOX” reaction in the model. The cloud pH will change in accordance with the additional sulfate produced via

“S(IV) + HOX” reaction, as cloud pH in the model is a function of sulfate and other ion concentrations.

Changes to the manuscript: Add “The magnitude of the potential decrease in cloud pH can only be obtained after adding the “S(IV) + HOX” reactions in the model, which will be done in a follow-up study.” to Line 422.

Reference

Faloona, I., Tan, D., Brune, W., Hurst, J., Barket, D., Couch, T.L., Shepson, P., Apel, E., Riemer, D., Thornberry, T. and Carroll, M.A.: Nighttime observations of anomalously high levels of hydroxyl radicals above a deciduous forest canopy, *J. Geophys. Res.: Atmospheres*, 106(D20), pp.24315-24333, 2001.

Fuchs, H., Hofzumahaus, A., Rohrer, F., Bohn, B., Brauers, T., Dorn, H.P., Häsel, R., Holland, F., Kaminski, M., Li, X. and Lu, K.: Experimental evidence for efficient hydroxyl radical regeneration in isoprene oxidation, *Nature Geoscience*, 6(12), pp.1023-1026, 2013

Lu, K. D., Rohrer, F., Holland, F., Fuchs, H., Brauers, T., Oebel, A., Dlugi, R., Hu, M., Li, X., Lou, S. R., Shao, M., Zhu, T., Wahner, A., Zhang, Y. H., and Hofzumahaus, A.: Nighttime observation and chemistry of HO_x in the Pearl River Delta and Beijing in summer 2006, *Atmos. Chem. Phys.*, 14, 4979-4999, doi:10.5194/acp-14-4979-2014, 2014.

Isotopic constraints on the role of hypohalous acids in sulfate aerosol formation in the remote marine boundary layer

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Abstract. Sulfate is an important component of global atmospheric aerosol, and has partially compensated for greenhouse gas-induced warming during the industrial period. The magnitude of direct and indirect radiative forcing of aerosols since preindustrial time is a large uncertainty in climate models, which has been attributed largely to uncertainties in the preindustrial environment. Here, we report observations of the oxygen isotopic composition ($\Delta^{17}\text{O}$) of sulfate aerosol collected in the remote marine boundary layer (MBL) in spring and summer in order to evaluate sulfate production mechanisms in pristine-like environments. Model-aided analysis of the observations suggests that 33-50% of sulfate in the MBL is formed via oxidation by hypohalous acids ($\text{HOX}=\text{HOBr}+\text{HOCl}$), a production mechanism typically excluded in large scale models due to uncertainties in the reaction rates, which are due mainly to uncertainties in reactive halogen concentrations. Based on the estimated fraction of sulfate formed via HOX oxidation, we further estimate that daily-averaged HOX mixing ratios on the order of 0.01-0.1 parts per trillion (ppt=pmol/mol) in the remote MBL during spring and summer are sufficient to explain the observations.

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1 Introduction

30 Large uncertainties in estimates of aerosol radiative forcing, especially those induced by sulfate aerosol and its interaction with clouds, significantly impede the progress of constraining the magnitude of anthropogenic radiative forcing since preindustrial times (Myhre et al., 2013). The aerosol radiative forcing uncertainties are attributed in large part to our poor understanding of the abundance of natural aerosols, especially sulfate aerosol in the marine boundary layer (MBL) (Carslaw et al., 2013) that is mainly produced from the oxidation of dimethylsulfide (DMS) emitted from oceanic phytoplankton (Bates et al., 1992). The radiative effects of sulfate involve scattering of solar radiation and modification of cloud properties (Haywood and Boucher, 2000). Determining the magnitude of these radiative effects requires in part understanding of sulfate formation mechanisms. Only sulfate formed via gas-phase oxidation can nucleate new particles (Kerminen et al., 2010; Kulmala et al., 2000) with implications for particle number density. Sulfate formed in the aqueous phase impacts particle growth rates in clouds with implications for aerosol size distribution (Lelieveld and Heintzenberg, 1992).

40 In the MBL, due to the high solubility and fast aqueous-phase oxidation of SO_2 , the main sulfate production mechanisms are thought to be in-cloud oxidation of dissolved SO_2 ($\text{S(IV)} = \text{SO}_2 \cdot \text{H}_2\text{O} + \text{HSO}_3^- + \text{SO}_3^{2-}$) by hydrogen peroxide (H_2O_2) and ozone (O_3) (Faloona, 2009; Alexander et al., 2012). In addition to sulfate formation in clouds, MBL sulfate formation can occur via oxidation of SO_2 by OH in the gas phase (Stockwell and Calvert, 1983) and on the surface of sea-salt aerosols in the presence of O_3 (Sievering et al., 2004; Alexander et al., 2005). Other sulfate production mechanisms that are important in specific environments, such as metal-catalyzed oxidation of S(IV) by O_2 (Alexander et al., 2009; Harris et al., 2013; 2014) and gas-phase oxidation of SO_2 by Criegee intermediates (Mauldin III et al., 2012; Pierce et al., 2013), are thought to be 45 minor in the MBL.

Modeling studies by Vogt et al. (1996) attributed a large part (60%) of aqueous-phase sulfate production in the MBL to oxidation of S(IV) by hypohalous acids ($\text{HOX} = \text{HOBr} + \text{HOCl}$); Further, von Glasow et al. (2002) evaluated the contribution of HOX to sulfate formation in both the cloud-free and cloudy MBL with a numerical one-dimensional model. Under both cloud-free and cloudy MBL conditions, about 30% of sulfate was formed via oxidation of S(IV) by HOX in the aqueous phase. Despite the potentially important role of HOX in sulfate formation in the MBL, the “S(IV)+HOX” reaction has not been included in most large-scale models of sulfur chemistry due to large uncertainties in (1) rate constants for reactions between HOX and HSO_3^- , (2) the Henry’s law constant for HOCl (H_{HOCl}) and HOBr (H_{HOBr}), and (3) concentrations of HOX in the MBL.

55 Laboratory experiments demonstrate that reactions between HOX and SO_3^{2-} occur via nucleophilic attack of SO_3^{2-} onto the X atom of HOX ($X = \text{Br}$ or Cl), followed by rapid hydrolysis of XSO_3^- (Yiin and Margerum, 1988; Fogelman et al., 1989; Troy and Margerum, 1991):





60 The rate constants for the “HOBr+SO₃²⁻” ($k_{\text{HOBr}+\text{SO}_3^{2-}}$) and “HOCl+SO₃²⁻” ($k_{\text{HOCl}+\text{SO}_3^{2-}}$) reactions are $5 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ and $7.6 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$, respectively (Fogelman et al., 1989; Troy and Margerum, 1991); In addition, Liu (2000) suggested that the reaction of HOBr with HSO₃⁻ follows a similar pathway as with SO₃²⁻, but with a lower reaction rate constant due to reduced nucleophilicity of HSO₃⁻ compared to SO₃²⁻:



They were unable to obtain $k_{\text{HOBr}+\text{HSO}_3^-}$ from their laboratory experiments due to interference from the unavoidable reaction of HOBr with Br⁻ in acidic solution. Based on their laboratory results, they suggested an upper limit for $k_{\text{HOBr}+\text{HSO}_3^-}$ ($< 3.2 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$). To our best knowledge, there is no laboratory experiment that determined the reaction rate constant of HOCl with HSO₃⁻ ($k_{\text{HOCl}+\text{HSO}_3^-}$). It is reasonable to assume that the reaction of HOCl with HSO₃⁻ follows a

70 similar pathway as the reaction of HOBr with HSO₃⁻ (Eqs. 3-4). Lack of knowledge of $k_{\text{HOBr}+\text{HSO}_3^-}$ and $k_{\text{HOCl}+\text{HSO}_3^-}$ leads to large uncertainties in calculations of the sulfate formation rate from HOX because HSO₃⁻ is the dominant S(IV) species (>93%) in clouds at typical cloud pH between 3 and 6 (Faloona, 2009).

Laboratory measurements of H_{HOCl} range from 470 to 910 M atm⁻¹ (Holzwarth et al., 1984; Hanson and Ravishankara, 1991; Blatchley III et al., 1992). Based on the aforementioned laboratory results, Huthwelker et al. (1995) suggested an expression for H_{HOCl} as a function of H₂SO₄ concentration and temperature. By assuming a temperature of 298.15 K and pure water, Sander et al. (2006) suggested $H_{\text{HOCl}} \approx 650 \text{ M atm}^{-1}$. Estimates of H_{HOBr} show a much larger range from 93 to 6100 M atm⁻¹ (McCoy et al., 1990; Blatchley III et al., 1992; Vogt et al., 1996; Frenzel et al., 1998; Sander et al., 2006; 2015). $H_{\text{HOBr}} = 93 \text{ M atm}^{-1}$ was assumed in the modeling studies by Vogt et al. (1996) and von Glasow et al. (2002), who simply estimated H_{HOBr} as 10% of the solubility constant of HOCl at 293 K from Huthwelker et al. (1995). Frenzel et al. (1998) estimated H_{HOBr} to be 6100 M atm⁻¹ using the Gibbs free energy of HOBr. Blatchley III et al. (1992) estimated H_{HOBr} to be at least twice the Henry’s law constant of HOCl that was measured in their laboratory experiments. Based on this relationship, Sander et al. (2006) extrapolated H_{HOBr} to be $\geq 1300 \text{ M atm}^{-1}$ using H_{HOCl} from Huthwelker et al. (1995). Only McCoy et al. (1990) measured H_{HOBr} from laboratory experiments and reported H_{HOBr} to be about 1900 M atm⁻¹.

85 Observations of HOX concentrations in the troposphere are sparse. Liao et al. (2012) made the first direct observation of HOBr in Alaska in spring 2009, and reported average daytime surface mixing ratios of about 10 ppt, consistent with the active bromine (HOBr+Br₂) mixing ratios measured by Neuman et al. (2010). HOBr mixing ratios were below their detection limit of 2 ppt at night (Liao et al., 2012). A recent aircraft campaign showed a daily-averaged HOBr mixing ratio of about 2 ppt along its flight track from 300 m to about 8000 m over the tropical Western Pacific during January and

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February 2014, with very small vertical gradients (Le Breton et al., 2016). The detection limit of HOBr in Le Breton et al. (2016) was about 0.2 ppt. The only direct observation of HOCl mixing ratio was made over the eastern tropical Atlantic at the surface during June 2009 and a large range from <5 ppt to 173 ppt was reported (Lawler et al., 2011). The detection limit of HOCl in Lawler et al. (2011) was 5 ppt.

The $\Delta^{17}\text{O}$ ($\approx \delta^{17}\text{O} - 0.52\delta^{18}\text{O}$) of sulfate is solely dependent upon the relative importance of the oxidants involved in its formation (Savarino et al., 2000), and thus provides an observational constraint for sulfate formation pathways (Lee and Thiemens, 2001; Lee et al., 2001; Jenkins and Bao, 2006; McCabe et al., 2006; Patris et al., 2007; Dominguez et al., 2008; Alexander et al., 2005; 2009; 2012). $\delta^{17}\text{O}$ or $\delta^{18}\text{O}$ is expressed as:

$$\delta^x\text{O} = \frac{R_{\text{SA}}^x}{R_{\text{VSMOW}}^x} - 1 \quad (5)$$

where R_{SA}^x is the $^{16}\text{O}/^{18}\text{O}$ ratio of the sample, R_{VSMOW}^x is the same ratio of Vienna Standard Mean Ocean Water (VSMOW) (Gonfiantini et al., 1993), and $x=17$ or 18. The $\Delta^{17}\text{O}$ value is expressed in unit of per mill (‰). Table 1 lists the $\Delta^{17}\text{O}$ of sulfate formed via different pathways. $\Delta^{17}\text{O}$ of sulfate produced from OH, H_2O_2 , and metal-catalyzed O_2 oxidation pathways are 0 ‰ (Dubey et al., 1997; Lyons, 2001), 0.7 ‰ (Savarino and Thiemens, 1999), and -0.09 ‰ (Barkan and Luz, 2005), respectively, which were discussed in detail in Alexander et al. (2005; 2009) and Sofen et al. (2011) and will not be repeated here. Primary anthropogenic sulfate has a $\Delta^{17}\text{O}$ of 0 ‰ (Lee et al., 2002). Sulfate produced from O_3 oxidation has a $\Delta^{17}\text{O}$ of 6.5 ‰, assuming $\Delta^{17}\text{O}(\text{O}_3) = 26$ ‰ (Vicars and Savarino, 2014). $\Delta^{17}\text{O}$ of sulfate produced from HOX oxidation has not been directly determined from laboratory experiments. Since HOX promotes sulfate formation via “ SO_3^{2-} +HOX” reactions by adding one oxygen atom from H_2O to sulfate instead of transferring its own oxygen atom (Fogelman et al., 1989; Troy and Margerum, 1991; Yiin and Margerum, 1988), the $\Delta^{17}\text{O}$ of sulfate produced from “ SO_3^{2-} +HOX” reactions is expected to have the same $\Delta^{17}\text{O}$ as water (0 ‰) (Gonfiantini et al., 1993). Liu (2000) suggests the reaction of HOBr with HSO_3^- follows a similar pathway as with SO_3^{2-} (Eqs. 1-2), resulting in $\Delta^{17}\text{O}$ of 0 ‰ for sulfate produced via this reaction. We assume that the reaction of HOCl with HSO_3^- follows a similar pathway as the reaction of HOBr with HSO_3^- (Eqs. 3-4) and produces sulfate with $\Delta^{17}\text{O}$ of 0 ‰. Based on these mechanistic studies, the $\Delta^{17}\text{O}$ of sulfate produced from HOX oxidation is expected to be 0 ‰.

Here, we report observations of $\Delta^{17}\text{O}$ of sulfate in atmospheric aerosols collected over a large spatial domain in the remote southern hemisphere MBL during spring and summer. We use these observations, combined with a global chemical transport model, to estimate the role of HOX in sulfate formation in the MBL.

2 Sampling and measurements

Aerosol samples were collected on quartz fiber filters (Whatman) using high volume air samplers located at the front of the ships from two ship cruises: (1) “Malaspina” as part of the Malaspina Circumnavigation Campaign on board of RV

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Hespérides (González-Gaya et al., 2014), and (2) “Xue-Long” as part of the 28th China Antarctic Research Expedition supported by the the Program of China Polar Environment Investigation and Assessment (Project No. CHINARE2011–2015) on board of the Xue-Long icebreaker. The quartz filters were pre-combusted at 450 °C for 8 hours and kept wrapped in aluminum foil and plastic ziplock bags before use. Most Xue-Long filters were changed every 48 hours while most Malaspina filters were changed every 24 hours. The sampler was connected to a wind direction vane to avoid contamination from the ship exhaust. After sampling, filters were kept wrapped in aluminum foil and plastic ziplock bags at -20 °C. Blank filters were processed as field samples. One-quarter of each Xue-Long filter and 1/8 of each Malaspina filter were sent to the University of Washington for isotope and concentration measurements. Our samples from the Malaspina campaign cover the track from Cádiz, Spain to Sydney, Australia via Rio de Janeiro, Brazil, Cape Town, South Africa and Perth, Australia between January 02 and March 23, 2011. The Xue-Long campaign starts from Shanghai China on November 4, 2011, traveling through West Australia, Zhong Shan station, Antarctica, South Argentina and back to Shanghai, China following the original route, ending on April 5, 2012. Figure 1 shows the 5-day back trajectories calculated from the NOAA HYSPLIT model for all sampling locations (<http://ready.arl.noaa.gov/HYSPLIT.php>), which gives a broad picture of the origins of air parcels along our sampling track. Most of the air parcels arriving at our sampling locations were over the ocean for the previous 5 days (\approx lifetime of sulfate (Chin et al., 2000)), which suggests that the sulfate collected was mainly formed in the MBL. However, observations of polycyclic aromatic hydrocarbons (PAHs) during the Malaspina campaign suggest that samples collected over the subtropical Indian Ocean and Atlantic Ocean might have continental influence (González-Gaya et al., 2014).

Aerosol ion concentrations (SO_4^{2-} , NO_3^- , Cl^- , NH_4^+ , Na^+ , K^+ , Mg^{2+} , Ca^{2+}) for filter samples from the Malaspina and Xue-Long campaigns were measured at South Dakota State University (USA) and University of Science and Technology of China (China), respectively. The analytical procedures for measurement of anions and cations using ion chromatography have been presented elsewhere (Cole-Dai et al., 1995; Jauhiainen et al., 1999). Typical instrumental analytical precision for all ions is better than 10% RSD (relative standard deviation) at the $\mu\text{g l}^{-1}$ level. Bromide aerosol concentrations ($[\text{Br}^-]$) were also measured for the Xue-Long samples (Supplementary data). There is no relationship between observed Br^- concentration and $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ nor our calculated $[\text{HOX}]_g$ (not shown) because factors such as aerosol pH, sunlight and oxidants play an important role in the acid-catalyzed formation of reactive halogens and removal of HOBr (Fickert et al., 1999; Schmidt et al., 2016). Similarly, there is no relationship between $[\text{Br}^-]$ and HOBr mixing ratios in the global modeling study by Schmidt et al. (2016) (not shown). Thus, $[\text{Br}^-]$ alone is not a good proxy for the “HOBr + S(IV)” reaction.

In the remote MBL, total sulfate consists of sea-salt sulfate (ssSO_4^{2-}) and non-sea-salt sulfate (nssSO_4^{2-}). ssSO_4^{2-} refers to primary sulfate emitted directly from sea water via the bursting of bubbles while nssSO_4^{2-} refers to secondary sulfate produced from oxidation of SO_2 . For the Xue-Long samples, the nssSO_4^{2-} fraction (f_{nss}) was calculated using the mass ratio of ($\text{ssSO}_4^{2-}/\text{Na}^+$) = 0.252 g/g in seawater (Millero et al., 2007). For the Malaspina samples, due to a sodium blank in the

160 quartz fiber filters, we calculated f_{nss} using the mass ratio of $(\text{ssSO}_4^{2-}/\text{Mg}^{2+}) = 2.115 \text{ g/g}$ in seawater (Millero et al., 2008). Only samples with f_{nss} larger than 30% were used in this study, to minimize the effect of uncertainty in the ssSO_4^{2-} fraction on calculations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$. In the end, 25 Malaspina and 42 Xue-Long samples were used in this study, for a total of 67 (out of 91) samples. The averaged f_{nss} is 0.57 ± 0.21 (1σ) and 0.74 ± 0.19 (1σ) for the 25 Malaspina and 42 Xue-Long samples, respectively.

165 The samples are divided into four categories (Fig. 2): (I) Southern Ocean, (II) Antarctic coast, (III) subtropical MBL and (IV) tropical coasts, based on their geographical location. The number of samples is 13, 18, 19 and 17 for Category I, II, III and IV, respectively.

$\Delta^{17}\text{O}$ of sulfate on the aerosol filter samples were measured using the pyrolysis method described in detail in Geng et al. (2013). Briefly, the sulfate on the filters was first dissolved in 18 M Ω water, purified using ion chromatography, and converted to Na_2SO_4 using ion exchange resin (AG 50 W-X8, 100–200 mesh, H⁺ form, Bio-Rad, Hercules, CA, USA). 30% H_2O_2 solution was added to remove organics, and Na_2SO_4 was then converted to Ag_2SO_4 using the ion exchange resin. The Ag_2SO_4 was dried to a solid in a quartz cup and each sample was placed in a zero-blank autosampler attached to the continuous-flow inlet of the isotope ratio mass spectrometer (IRMS). Each Ag_2SO_4 sample was individually dropped into a furnace (1000°C) under a continuous flow of helium (He) where it is pyrolyzed to form $\text{Ag}(\text{s}) + \text{SO}_2(\text{g}) + \text{O}_2(\text{g})$. The byproduct $\text{Ag}(\text{s})$ condenses on the walls of the quartz pyrolysis tube, while the byproduct $\text{SO}_2(\text{g})$ was removed from the He flow with a cryogenic trap at liquid nitrogen temperature ($\approx 77 \text{ K}$). The remaining product $\text{O}_2(\text{g})$ is carried along the He flow to the IRMS, where the abundance of ^{16}O , ^{17}O , and ^{18}O in O_2 was measured, and from which $\Delta^{17}\text{O}$ was calculated. 54 samples were measured in triplicate, 9 samples were measured in duplicate, and 4 samples were measured once. The precision of $\Delta^{17}\text{O}$ is typically better than $\pm 0.3\%$ based on replicate analysis of standards. The $\Delta^{17}\text{O}$ obtained from IRMS is the $\Delta^{17}\text{O}$ of total sulfate on the aerosol samples ($\Delta^{17}\text{O}(\text{SO}_4^{2-})$). $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ was calculated by dividing $\Delta^{17}\text{O}(\text{SO}_4^{2-})$ by f_{nss} , as $\Delta^{17}\text{O}(\text{ssSO}_4^{2-}) = 0 \%$ (Bao et al., 2000).

3 GEOS-Chem model

We use v9-02 of the GEOS-Chem global 3-D model (<http://www.geos-chem.org/>) of coupled oxidant-aerosol chemistry (Park et al., 2004) to simulate atmospheric sulfur chemistry and interpret our $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ observations. The model is driven by assimilated meteorological data from the NASA Goddard Earth Observing System (GEOS-5, <http://gmao.gsfc.nasa.gov>), which has a temporal resolution of 6 hours, with 3 hours for surface quantities and mixing depths. Model simulations for the analysis of the cruise data were performed at $2^\circ \times 2.5^\circ$ horizontal resolution and 47 vertical levels up to 0.01 hPa using GEOS-5 meteorology corresponding to the timing of sample collection, after spinning up the model for one year.

190 The sulfate simulations were conducted in aerosol-only mode that used archived monthly mean OH, NO₃, O₃ and total
nitrate concentrations and production and loss rates for H₂O₂ from the full-chemistry simulation as described in Park et al.
(2004). [A diurnal variation as a function of solar zenith angle is applied to OH concentrations and photolytic loss rates of
H₂O₂ in the model. NO₃ is set to be zero during daytime.](#) Sulfate in the model was produced from gas-phase oxidation of
SO_{2(g)} by OH, aqueous-phase oxidation of S(IV) by H₂O₂, O₃, and metal-catalyzed O₂ (Alexander et al., 2009), and
195 heterogeneous oxidation on sea-salt aerosols by O₃ (Alexander et al., 2005). [The parameterization of the metal-catalyzed
S\(IV\) oxidation pathway is described in detail in Alexander et al. \(2009\). The trace metals included are Fe and Mn, whose
oxidation states Fe\(III\) and Mn\(II\) catalyze S\(IV\) oxidation. Soil-derived Fe \(\[Fe\]_{soil}\) is scaled to modeled dust concentration
as 3.5% of total dust mass while soil-derived Mn \(\[Mn\]_{soil}\) is a factor of 50 lower than \[Fe\]_{soil}. Anthropogenic Mn \(\[Mn\]_{anthro}\)
is scaled as 1/300 of primary sulfate concentration while anthropogenic Fe \(\[Fe\]_{anthro}\) is 10 times that of \[Mn\]_{anthro}. We
200 assume that 50% of Mn is dissolved in cloud water as Mn\(II\) oxidation state. For Fe, we assume that 10% of \[Fe\]_{anthro} and 1%
of \[Fe\]_{soil} is dissolved in cloud water. 10% of the dissolved Fe is in Fe\(III\) oxidation state during daytime and 90% at night.](#)
Primary anthropogenic emissions of sulfate are 3.5% of total anthropogenic sulfur emissions in Europe, an average of 1.5%
in North America and 2.1% elsewhere. The anthropogenic emission inventories used in this study is the global emission
inventory EDGAR v3 (Olivier et al., 2001), supplemented by regional inventories such as STREETS (Streets et al., 2006),
205 EMEP (Vestreng and Klein, 2002), and NEI2005 (Van Donkelaar et al., 2015). The oceanic DMS inventory is from Kettle et
al. (1999). Sulfate formed from each oxidation pathway was treated as a separate “tracer” in the model with a corresponding
 $\Delta^{17}\text{O}$ value as shown in Table 1. Primary anthropogenic sulfate has a $\Delta^{17}\text{O}$ of 0 ‰ (Lee et al., 2002). [The model calculates
 \$\Delta^{17}\text{O}\$ of bulk sulfate in the grid box \(\$\Delta^{17}\text{O}_{\text{mod}}\(\text{nssSO}_4^{2-}\)\$ \) based on the relative importance of each sulfate production
mechanism for total sulfate abundance. This is then compared to the \$\Delta^{17}\text{O}\$ measurement of bulk sulfate collected on aerosol
210 filters \(\$\Delta^{17}\text{O}_{\text{obs}}\(\text{nssSO}_4^{2-}\)\$ \) to investigate sulfate formation mechanisms in the MBL \(Sect. 4.2\).](#)

For pH-dependent S(IV) partitioning, bulk cloud water pH is calculated as described in Alexander et al. (2012). Large-
scale models such as GEOS-Chem calculate the average chemistry of bulk cloud water rather than the chemistry of
individual cloud droplets. This approach has been shown to significantly underestimate sulfate formation via oxidation of
SO₃²⁻ by O₃ by underestimating the fraction of S(IV) present as SO₃²⁻ (Hegg et al., 1992; O’Dowd et al., 2000; Roelofs, 1993;
215 Yuen et al., 1996; Fahey and Pandis, 2003). We use the Fahey and Pandis (2003) algorithm and the Yuen et al. (1996)
parameterization in GEOS-Chem to account for the effect of heterogeneity in cloud drop pH on S(IV) partitioning as
described in Alexander et al. (2012).

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4 Results

4.1 Observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ and nssSO_4^{2-} concentration

Except for one sample near the coast of China with a relatively high concentration of nssSO_4^{2-} ($[\text{SO}_4^{2-}]_{\text{nss}}$) ($7.4 \mu\text{g m}^{-3}$), the $[\text{SO}_4^{2-}]_{\text{nss}}$ observations vary from 0.2 to $3.5 \mu\text{g m}^{-3}$, with an average of 1.2 ± 0.8 (1σ) $\mu\text{g m}^{-3}$. Averaged $[\text{SO}_4^{2-}]_{\text{nss}}$ is 1.4 ± 0.8 (1σ) $\mu\text{g m}^{-3}$, 1.4 ± 0.7 (1σ) $\mu\text{g m}^{-3}$, 0.9 ± 0.5 (1σ) $\mu\text{g m}^{-3}$ and 1.3 ± 0.9 (1σ) $\mu\text{g m}^{-3}$ for samples in Category I, II, III and IV, respectively (Table 2). A latitudinal gradient of $[\text{SO}_4^{2-}]_{\text{nss}}$ is found in our data, where averaged $[\text{SO}_4^{2-}]_{\text{nss}}$ between 50°S and 70°S is 50% higher than samples between 20°S and 40°S (1.5 versus $1.0 \mu\text{g m}^{-3}$). The difference is significant at the 95% confidence level.

Figure 2 shows the observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ ($\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$) and nssSO_4^{2-} concentration ($[\text{SO}_4^{2-}]_{\text{nss}}$) at each sampling location. $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ values range from 0.0 to 1.6 ‰, with an average of 0.7 ± 0.4 (1σ) ‰. Averaged $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ (Table 2) is 0.5 ± 0.3 (1σ) ‰, 0.7 ± 0.4 (1σ) ‰, 0.8 ± 0.4 (1σ) ‰ and 0.8 ± 0.4 (1σ) ‰ for samples in Category I, II, III and IV, respectively. The analytical error in $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ is estimated by calculating the standard deviation (1σ) of the multiple measurements of each sample, which range from ± 0.0 - 0.4 ‰ with an average of ± 0.1 ‰. Though $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ in Category I (over Southern Ocean) is slightly lower than those in other categories, the < 0.3 ‰ difference between each category is generally smaller than the measurement uncertainty estimated from replicate analysis of standards.

Only sulfate formed via H_2O_2 and O_3 oxidation has a positive $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$, with H_2O_2 oxidation leading to $\Delta^{17}\text{O}(\text{nssSO}_4^{2-}) = 0.7$ ‰ and O_3 oxidation leading to $\Delta^{17}\text{O}(\text{nssSO}_4^{2-}) = 6.5$ ‰. We can calculate the maximum contribution from “S(IV)+ O_3 ” ($f_{\text{O}_3, \text{max}}$) for each sulfate sample by assuming no contribution from H_2O_2 (i.e. all of the aqueous-phase S(IV) oxidation occurs via O_3 oxidation):

$$f_{\text{O}_3, \text{max}} = \frac{\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})}{\Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{O}_3}} \quad (6)$$

where $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{O}_3} = 6.5$ ‰. This yields $f_{\text{O}_3, \text{max}}$ ranging from 0.00 to 0.26 with an average of 0.12 ± 0.06 (1σ) for all samples. Averaged $f_{\text{O}_3, \text{max}}$ is 0.08 ± 0.05 (1σ), 0.11 ± 0.06 (1σ), 0.13 ± 0.07 (1σ) and 0.13 ± 0.06 (1σ) for samples in Category I, II, III and IV, respectively (Table 2). For samples with $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ larger than 0.7 ‰ (36 samples), we can calculate the minimum O_3 contribution ($f_{\text{O}_3, \text{min}}$) by assuming that H_2O_2 is the only other oxidation pathway (i.e., no significant contribution from OH and HOX oxidation):

$$f_{\text{O}_3, \text{min}} = \frac{\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-}) - \Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{H}_2\text{O}_2}}{\Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{O}_3} - \Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{H}_2\text{O}_2}} \quad (7)$$

where $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{H}_2\text{O}_2} = 0.7 \text{ ‰}$. For samples with $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ smaller than 0.7 ‰ (31 samples), $f_{\text{O}_3,\text{min}}$ is 0. We obtain $f_{\text{O}_3,\text{min}}$ ranging from 0.00 to 0.16 with an average of 0.03 ± 0.04 (1σ) among all samples. Averaged $f_{\text{O}_3,\text{min}}$ is 0.01 ± 0.02 (1σ), 0.03 ± 0.04 (1σ), 0.04 ± 0.05 (1σ) and 0.03 ± 0.04 (1σ) for samples in Category I, II, III and IV, respectively (Table 2).

250 4.2 Comparison of modeled versus observed $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$

Figure 3a shows the comparison between modeled and observed $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ ($\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ vs. $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$) for the standard model run (as described in section 3). $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ represents the daily mean in the first vertical model level (below ≈ 100 m) at each of our sampling locations. The range of $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ is $0.7 \sim 5.6 \text{ ‰}$, overestimating the observations on average by a factor of 2.5. Averaged $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ is 3.6 ± 1.1 (1σ) ‰, 1.1 ± 0.3 (1σ) ‰, 1.9 ± 1.0 (1σ) ‰ and 1.2 ± 0.4 (1σ) ‰ for samples in Category I, II, III and IV, respectively. $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ overestimates $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ in all categories (Table 2). The discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ is most evident for samples in Category I (Southern Ocean), for which the model predicts 48% of sulfate is formed via O_3 pathway, compared to 1~8% estimated from the observations alone.

The model calculated fractional contributions to the sulfate burden from each oxidant, averaged over all samples, is $f_{\text{mod,OH}} = 0.20 \pm 0.14$ (1σ), $f_{\text{mod,H}_2\text{O}_2} = 0.57 \pm 0.15$ (1σ), $f_{\text{mod,O}_3} = 0.20 \pm 0.19$ (1σ) and $f_{\text{mod,het}} = 0.02 \pm 0.03$ (1σ), where $f_{\text{mod,OH}}$, $f_{\text{mod,H}_2\text{O}_2}$, $f_{\text{mod,O}_3}$ and $f_{\text{mod,het}}$ represents the fractional contribution of gas-phase OH oxidation, in-cloud H_2O_2 oxidation, in-cloud O_3 oxidation and heterogeneous oxidation by O_3 on the surface of sea salt aerosol to the total sulfate burden, respectively. The corresponding fractional contributions for samples in different categories are shown in Table 2. $f_{\text{mod,O}_3}$ is largest in Category I (0.48) and smallest in Category II and IV (0.10) while $f_{\text{mod,OH}}$ is largest in Category IV (0.30), and smallest in Category I (0.04). Sulfate formation from in-cloud metal-catalyzed oxidation by O_2 and direct emission of anthropogenic sulfate contribute less than 1% of total sulfate in our samples and thus will be neglected in the discussion below.

Based on the modeled fractional contributions to the sulfate burden from each oxidant except HOX, and knowledge about reaction rate constants of “HOX + S(IV)” reactions and Henry’s law constants of HOX, we calculate the amount of HOX needed to explain the discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ in Sect. 5.2.4.

270 5 Discussion

5.1 Observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ and nssSO_4^{2-} concentration

Our observations of nssSO_4^{2-} concentration ($0.2 \sim 3.5 \mu\text{g m}^{-3}$) are consistent with those ($< 2 \mu\text{g m}^{-3}$) over the Southern Ocean measured by Sievering et al. (2004). Other published data for samples with air originating from the remote Atlantic Ocean showed a $[\text{SO}_4^{2-}]_{\text{nss}}$ between 0.9 and $4.5 \mu\text{g m}^{-3}$ (Alexander et al., 2012), consistent with our observations. Higher observed $[\text{SO}_4^{2-}]_{\text{nss}}$ between $50^\circ\text{S} - 70^\circ\text{S}$ compared to $20^\circ\text{S} - 40^\circ\text{S}$ could be due to a higher DMS emission flux over $50^\circ - 70^\circ\text{S}$ (Boucher et al., 2003).

Previous studies have suggested a large contribution to sulfate formation from O₃ oxidation in the MBL as the aqueous-phase reaction between S(IV) and O₃ is very fast at pH > 5 (Chameides and Stelson, 1992; Sievering et al, 1991; 2004; O'Dowd et al., 2000; Alexander et al., 2012). These studies did not consider the HOX mechanism due to the large
280 uncertainty in the reaction rates. As the reaction of HOX with SO₃²⁻ is also fast ($k_{\text{HOBr}+\text{SO}_3^{2-}} = 5 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ and $k_{\text{HOCl}+\text{SO}_3^{2-}} = 7.6 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$), cloud pH > 5 will promote rapid aqueous-phase sulfate formation by HOX in addition to O₃. A large contribution from O₃ will yield a high $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ value (6.5 ‰), but most samples in this study have low $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ values (0.7±0.4 ‰). Thus, our results argue against a significant role of sulfate formation via O₃ oxidation in remote MBL. Indeed, our calculated O₃ contribution range ($f_{\text{O}_3,\text{min}}=0.03, f_{\text{O}_3,\text{max}}=0.12$) is more consistent with that reported by
285 von Glasow et al. (2002), who did consider HOX. Their simulations of sulfate production in the MBL yielded a calculated O₃ contribution of 2-8% to the total sulfate production, while HOX contributed about 30%.

In comparison, Alexander et al. (2012) showed observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ of 1.1-1.4 ‰ for samples with back trajectories over the Iberian Peninsula during summer and 2.2-7.3 ‰ for samples with back trajectories over the Atlantic Ocean during winter. Their results suggested sulfate formation via HOX oxidation is not significant over subtropical
290 northeast Atlantic during winter but potentially important in the more polluted coastal location of the Iberian Peninsula during summer. Our samples show lower $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ than Alexander et al. (2012) in general, but were collected at different locations and during different seasons.

5.2 Comparison of modeled versus observed $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$

As shown in Fig 3a, the standard model significantly overestimates observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$. This could be caused by
295 the model (1) overestimating (O₃) or underestimating (OH, H₂O₂) oxidant abundances, (2) overestimating the amount of clouds, (3) overestimating the pH of clouds, or (4) neglecting sulfate formation from HOX oxidation. In this section, we examine each of these possibilities.

5.2.1 Oxidant sensitivity simulations

To investigate the impact of model biases in oxidant concentrations on calculated $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$, we perform three
300 sensitivity runs by (1) doubling OH concentrations, (2) halving O₃ concentration and (3) doubling H₂O₂ concentration everywhere in the model. Figure 3b-d show $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ versus $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ for these three sensitivity runs. The discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ is not reconciled by changing the oxidant concentrations in the model. On average, $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ changes to 1.9±1.3 ‰, 1.8±1.2 ‰ and 1.4±0.9 ‰ for doubling OH concentration, halving O₃ concentration and doubling H₂O₂ concentration, respectively, compared to 1.8±1.2 ‰ for the standard run.

305 Doubling modeled OH concentrations results in an increase in the mass fraction of sulfate formed via gas-phase oxidation by OH ($f_{\text{mod,OH}}$) from 0.20 to 0.27 and a decrease via aqueous-phase oxidation by H₂O₂ and O₃ ($f_{\text{mod,aq}}=f_{\text{mod,H}_2\text{O}_2}+f_{\text{mod,O}_3}$) from

0.77 to 0.70 (Table 3a). This would be expected to result in lower $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ as long as the relative importance of H_2O_2 and O_3 in the aqueous phase remains constant. However, doubling OH results in a small increase in $f_{\text{mod},\text{O}_3}$ from 0.20 to 0.22 (Table 3a). The small increase in $f_{\text{mod},\text{O}_3}$ occurs because of the reduction in the aqueous-phase sulfate production rate, which is caused by lower S(IV) due to faster removal of SO_2 by OH in the gas phase. A reduction in the aqueous-phase sulfate formation rate results in higher cloud-water pH, which increases the fraction of sulfate produced from O_3 oxidation, even though the total aqueous-phase sulfate production decreases. Thus, doubling OH concentrations has an insignificant effect on $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$. The nighttime OH concentrations observed in forests and urban areas (Faloona et al., 2001; Lu et al., 2014) should also have insignificant effects on our model results as most of our samples are collected in the remote MBL where isoprene and VOCs abundances are low. Our sensitivity study with doubled OH suggests additional nighttime sources of OH would not resolve the modeled overestimate of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ observations.

Halving modeled O_3 concentrations results in a small decrease in $f_{\text{mod},\text{O}_3}$ from 0.20 to 0.19 and a change of less than 0.01 in $f_{\text{mod},\text{aq}}$ (Table 3b). The decrease in $f_{\text{mod},\text{O}_3}$ is small because $f_{\text{mod},\text{O}_3}$ is mainly regulated by the cloud pH rather than O_3 abundance. In other words, the sulfate burden from O_3 oxidation is limited by concentration of SO_3^{2-} , not O_3 . As a result, halving O_3 has an insignificant effect on $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$.

Doubling modeled H_2O_2 concentrations results in an increase in $f_{\text{mod},\text{H}_2\text{O}_2}$ from 0.57 to 0.66, a decrease in $f_{\text{mod},\text{O}_3}$ from 0.20 to 0.14 (Table 3c) and an increase in $f_{\text{mod},\text{aq}}$ from 0.77 to 0.80. The increase in $f_{\text{mod},\text{H}_2\text{O}_2}$ causes an increase in $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ of less than 0.1 ‰, which is a small effect compared to the change in $f_{\text{mod},\text{O}_3}$ that results in a decrease in $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ of 0.4 ‰. The decrease in $f_{\text{mod},\text{O}_3}$ is caused by the decrease in cloud pH due to higher aqueous-phase sulfate production rate. Although doubling H_2O_2 results in a decrease in $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ of 0.4 ‰ on average, it is still too small to reconcile the 1.1 ‰ discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$.

5.2.2 Cloud fraction sensitivity simulations

To assess the uncertainty regarding the modeled cloud amount, we perform a sensitivity study by halving the cloud fraction in the model. As in-cloud S(IV) oxidation by H_2O_2 and O_3 produces sulfate with $\Delta^{17}\text{O} > 0$, a decrease in cloud fraction can potentially reduce $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ by reducing in-cloud sulfate formation.

Figure 3e shows the comparison between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ for the sensitivity simulation where cloud fraction is halved. The discrepancy between the model and observations is similar to that in the standard run. Averaged $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ for samples in different categories are shown in Table 3d. Compared to the standard run, averaged $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ for all samples changes less than 0.1 ‰. Halving cloud fraction results in a decrease in $f_{\text{mod},\text{aq}}$ from 0.77 to 0.70 and an increase in $f_{\text{mod},\text{OH}}$ from 0.20 to 0.26. The change of $f_{\text{mod},\text{O}_3}$ is less than 0.01. A large decrease of $f_{\text{mod},\text{O}_3}$ is not observed by halving cloud fraction because lowering aqueous-phase sulfate production rates results in higher cloud pH,

shifting the relative importance of H₂O₂ and O₃ oxidation in the aqueous phase. Halving the cloud amount thus does not resolve the discrepancy between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$.

5.2.3 Cloud pH sensitivity simulations

340 Previous work has shown that bulk cloud models tend to underestimate sulfate formed via the O₃ pathway by underestimating pH and thus underestimating the fraction of S(IV) that is in the form of SO₃²⁻. Yuen et al. (1996) developed a parameterization to correct for the underestimate in cloud pH by comparing a bulk cloud model with a cloud model that resolves the heterogeneity in cloud chemistry. The uncertainty in this parameterization, and thus the degree to which it might result in an overestimate of the contribution of O₃ to in-cloud sulfate formation in GEOS-Chem is difficult to assess.

345 We perform a sensitivity study which neglects heterogeneity in cloud chemistry by turning off the Yuen et al. (1996) parameterization. By using only bulk cloud pH calculations, this is effectively a lower limit for cloud pH, and thus is a lower limit for the contribution of O₃ to in-cloud sulfate formation in the model in the absence of HOX.

Figure 3f shows the comparison between $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ for this low cloud pH simulation. Overall, $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ matches $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ much better than the standard run, especially for samples over Southern Ocean. $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ ranges from 0.4 ‰ to 2.4 ‰, with an average of 1.1 ± 0.5 (1 σ) ‰. In comparison, $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ ranges from 0.0 ‰ to 1.6 ‰, with an average of 0.7 ± 0.4 (1 σ) ‰. The difference between averaged $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ and $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ (0.4 ‰) is just slightly larger than the measurement uncertainty (± 0.3 ‰). Compared to the standard run, $f_{\text{mod},\text{O}_3}$ (0.07 vs. 0.20) is much smaller. $f_{\text{mod},\text{O}_3}$ of 0.07 is within the O₃ contribution fraction range constrained by the observations ($f_{\text{min},\text{O}_3}=0.03, f_{\text{max},\text{O}_3}=0.12$). The decrease of $f_{\text{mod},\text{O}_3}$ is the main reason for the drop in $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$. For samples over Southern Ocean, $f_{\text{mod},\text{O}_3}$ decreases from 0.48 to 0.08 and $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ decreases from 3.6 ‰ to 1.3 ‰ correspondingly (Table 3e).

Though averaged $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ is not much larger than $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ in the low cloud pH sensitivity study, the model does a poor job matching observations for samples with $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-}) < 0.7$ ‰. 25 out of 31 samples with $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ smaller than 0.7 ‰ show that $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ overestimates $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ by more than 50% (0.9 ‰), which indicates the model's inability to produce sulfate with low $\Delta^{17}\text{O}$ even while underestimating cloud pH. The majority of the discrepancy occurs for samples in the Southern Ocean (Category I) and subtropical MBL (Category III).

5.2.4 Contribution of HOX oxidation to sulfate formation

We can estimate the fractional contribution of HOX ($f_{\text{obs},\text{HOX}}$) to total sulfate abundance necessary to explain the low $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ by comparing modeled $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ with observations. $f_{\text{obs},\text{HOX}}$ is calculated as shown below:

$$365 \quad f_{\text{obs},\text{HOX}} = 1 - \frac{\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})} \quad (8)$$

The derivation of Eq. (8) is described in the Appendix A. Calculating $f_{\text{obs,HOX}}$ using Eq. (8) may overestimate $f_{\text{obs,HOX}}$, as it assumes that the addition of “S(IV)+HOX” will not impact cloud pH. We estimate that this assumption overestimates calculation of $f_{\text{obs,HOX}}$ by about 15% (see Appendix A).

370 For $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ in Eq. (8), we use results from both the standard run in Sect. 4.2 and the low cloud pH run in Sect. 5.2.3 to place bounds on $f_{\text{obs,HOX}}$, using the low cloud pH sensitivity study as a lower limit for $f_{\text{obs,HOX}}$. We assume $f_{\text{obs,HOX}}=0$ when $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-}) > \Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ (12 samples in the standard run and 22 samples in the low cloud pH run). $f_{\text{obs,HOX}}$ for each sample is shown in Fig. 4 for both runs. The averaged $f_{\text{obs,HOX}}$ is shown for samples in each category in Table 4a (standard run) and 4b (low cloud pH run). Among all samples, averaged $f_{\text{obs,HOX}}$ is 0.50 ± 0.33 for the standard run and 0.33 ± 0.32 for the low cloud pH run. $f_{\text{obs,HOX}}$ is largest for samples over the Southern Ocean (0.58-0.84 on average), 375 followed by the subtropical MBL (0.36-0.47 on average). $f_{\text{obs,HOX}}$ is lowest for samples collected near tropical coasts in the standard run (0.35) and near the Antarctic coast in the low cloud pH run (0.18).

We can estimate the concentration of HOX needed to achieve $f_{\text{obs,HOX}}$ using Eq. (9) below. The derivation of Eq. (9) is described in the Appendix A.

$$[\text{HOX}]_{\text{aq}} = \frac{f_{\text{obs,HOX}}}{\frac{k_{\text{HOX}+\text{HSO}_3^-} f_{\text{obs,H}_2\text{O}_2} + k_{\text{HOX}+\text{SO}_3^{2-}} f_{\text{obs,O}_3}}{k_{\text{H}_2\text{O}_2+\text{HSO}_3^-} [\text{H}_2\text{O}_2]_{\text{aq}} + k_{\text{O}_3+\text{SO}_3^{2-}} [\text{O}_3]_{\text{aq}}}} \quad (9)$$

380 where $k_{\text{H}_2\text{O}_2+\text{HSO}_3^-} = 2.4 \times 10^3 \text{ M}^{-1}\text{s}^{-1}$ at pH = 4.5 and $k_{\text{O}_3+\text{SO}_3^{2-}} = 1.5 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$. $k_{\text{HOBr}+\text{SO}_3^{2-}} = 5 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ and $k_{\text{HOCl}+\text{SO}_3^{2-}} = 7.6 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$ are from Fogelman et al. (1989) and Troy and Margerum (1991), respectively. We assume $k_{\text{HOX}+\text{SO}_3^{2-}} = 2.9 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ as the average of $k_{\text{HOBr}+\text{SO}_3^{2-}}$ ($= 5 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$, Troy and Margerum (1991)) and $k_{\text{HOCl}+\text{SO}_3^{2-}}$ ($= 7.6 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$, Fogelman et al. (1989)). We assume $k_{\text{HOX}+\text{HSO}_3^-} = 2.0 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$ which is the average of $k_{\text{HOBr}+\text{HSO}_3^-}$ ($= 3.2 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$, upper limit from Liu (2000)) and $k_{\text{HOCl}+\text{HSO}_3^-}$ ($= k_{\text{HOCl}+\text{SO}_3^{2-}}$, as assumed by Vogt et al. 385 (1996) and von Glasow et al. (2002)). $f_{\text{obs,HOX}}$ is calculated from Eq. (8), and $f_{\text{obs,H}_2\text{O}_2}$ and $f_{\text{obs,O}_3}$ are calculated from Eqs. (A9 and A14) using the same assumption as $f_{\text{obs,HOX}}$ (see Appendix A). $[\text{H}_2\text{O}_2]_{\text{aq}}$ and $[\text{O}_3]_{\text{aq}}$ are obtained from the model, and range from 2-172 μM and 113-463 pM, respectively. The range and median value of $[\text{HOX}]_{\text{aq}}$ for samples in different categories are shown in Table 4. $[\text{HOX}]_{\text{aq}}$ needed to explain $f_{\text{obs,HOX}}$ is on the order of 100 pM and 10 pM for the standard run and low cloud pH run, respectively. By assuming the Henry’s law constant H_{HOX} of 975 M atm^{-1} (average between H_{HOCl} and 390 H_{HOBr} from Huthwelker et al. (1995) and Sander et al. (2006)), the daily-averaged gas-phase $[\text{HOX}]_{\text{g}}$ mixing ratio $[\text{HOX}]_{\text{g}}$ is calculated and shown in Table 4. Due to the low solubility of HOX, under typical atmospheric conditions, more than 99% of total HOX is in the gas phase. Daily-averaged $[\text{HOX}]_{\text{g}}$ is on the order of 0.1 ppt and 0.01 ppt when using $f_{\text{obs,HOX}}$ from the standard run and low cloud pH run, respectively. Thus, a gas-phase HOX mixing ratio of ≈ 0.1 ppt or higher would be sufficient to explain the observed $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ of our samples. Uncertainties in our calculated $[\text{HOX}]_{\text{g}}$ can originate from

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(1) uncertainties in the rate constant for reaction between HOX and HSO_3^- ($k_{\text{HOX}+\text{HSO}_3^-}$), (2) uncertainties in the Henry's law constant for HOX (H_{HOX}) and (3) the efficiency of reactive uptake of gas-phase HOX onto cloud droplets that is not accounted for in our assumption of equilibrium of HOX between the gas and aqueous phase.

In comparison, a box-modeling study by Vogt et al. (1996) estimated that daytime-averaged $[\text{HOX}]_g$ on the order of 10 ppt is needed to achieve a similar fraction ($\approx 60\%$) of sulfate formed via HOX oxidation. The difference in $[\text{HOX}]_g$ between our study and Vogt et al. (1996) is caused by several factors. First, H_{HOBr} in our calculations is an order of magnitude larger than that in Vogt et al. (1996), so that our calculations require an order of magnitude lower $[\text{HOBr}]_g$ to produce the same aqueous-phase concentration. Second, H_2O_2 and O_3 mixing ratios in our calculations (Eq. 9) are lower than those in Vogt et al. (1996) (0.6 ppt vs. 0.8 ppt for $[\text{H}_2\text{O}_2]$ and 18 ppt vs. 40 ppt for $[\text{O}_3]$), so that higher $[\text{HOBr}]_g$ is needed in Vogt et al. (1996) to compete with S(IV) oxidation by H_2O_2 and O_3 . Third, in our simple calculation we assume equilibrium of HOX between the gas- and aqueous-phase while Vogt et al. (1996) considers all mass transfer limitations. Higher $[\text{HOX}]_g$ will be calculated if diffusion and subsequent mass accommodation of gas-phase HOX onto the cloud droplets is not fast enough to compensate for the loss of HOX from aqueous-phase chemistry. Fourth, $[\text{HOX}]_g$ on the order of 0.1 ppt calculated in our study is a daily-averaged concentration while $[\text{HOX}]_g$ on the order of 10 ppt in Vogt et al. (1996) is a daytime-averaged concentration, and hence are not directly comparable. Vogt et al. (1996) has shown that nighttime-averaged $[\text{HOX}]_g$ is about two orders of magnitude lower than daytime-averaged $[\text{HOX}]_g$.

The daily-averaged HOBr mixing ratio over Southern Ocean ($40^\circ\sim 64^\circ\text{S}$, below 100 m) modeled by Schmidt et al. (2016) varied from 0.1-0.3 ppt to 0.2-0.4 ppt for simulations without and with debromination from acidic sea salt aerosol, respectively. The HOX mixing ratio on the order of 0.1 ppt calculated from $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ of our samples using standard run results is consistent with that obtained in Schmidt et al. (2016).

Comparison of our calculated daily-averaged HOX mixing ratios with observations is difficult, as HOX is expected to show significant diurnal variability (on the order of 2 ppt), with mixing ratios near zero at night and peaking at around noon (von Glasow et al., 2002). A daytime-averaged HOBr mixing ratio of about 10 ppt was observed by Liao et al. (2012) at Alaska in spring, which is about 2~3 orders of magnitude higher than our calculated daily-averaged $[\text{HOX}]_g$. The nighttime HOBr mixing ratio in their study was below the detection limit of about 2 ppt. A more recent study that measured HOBr mixing ratios along the flight tracks over the tropical West Pacific during January and February 2014 showed a daily-averaged HOBr mixing ratio of about 2 ppt, with a lower detection limit of about 0.2 ppt (Le Breton et al., 2016). This is still much higher than our calculated daily-averaged $[\text{HOX}]_g$, but it is likely that HOBr mixing ratios could vary significantly with sampling locations and sampling time (Schmidt et al., 2016). Field campaigns of HOX measurements are necessary to assess our calculated HOX mixing ratios over our sampling regions.

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6 Conclusion

440 This study uses a combination of observations and modeling of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ to quantify the role of HOX (= HOBr + HOCl) in sulfate formation in the remote MBL. Samples collected over a wide spatial range in the MBL during spring and summer show low $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ ($0.7 \pm 0.4 \%$), which suggests only 3%~12% of sulfate is formed via O_3 oxidation. The standard model run overestimates $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ by about a factor of 2.5 on average because it overestimates the amount of sulfate formed by O_3 in the aqueous phase. This discrepancy could not be resolved by either varying oxidant concentrations, halving cloud amount or using a lower limit for cloud pH in the model. Our calculations suggest that the discrepancy can be explained with a fractional contribution of sulfate abundance formed by HOX ranging from 33~50% over the entire area sampled, with the highest fraction (58~84%) in the Southern Ocean MBL. A daily-averaged gas-phase HOX mixing ratio of ≈ 0.1 ppt or higher would be sufficient to explain the observed $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ of our samples. This study provides the first observational constraint on the role of hypohalous acids in sulfate aerosol formation in the MBL.

445

450 Future studies will implement the “S(IV) + HOX” reaction into GEOS-Chem to investigate the impacts of this reaction on the global sulfur budget and possible feedbacks on acid-catalyzed reactive halogen production.

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455

Appendix A

460 Hitherto, there is no observational constraint on HOX mixing ratios in the mid- and high-latitude remote MBL. Models have suggested a large range of HOX mixing ratios on the order of 0.1 ppt (Schmidt et al., 2016) to 10 ppt (Vogt et al., 1996) over these regions. Here we quantify the daily-averaged HOX mixing ratio indirectly from observed and modeled $\Delta^{17}\text{O}$ of sulfate. First, we calculate the fractional contribution of the HOX oxidation pathway ($f_{\text{obs,HOX}}$) to sulfate abundance in our samples. Then, we calculate the HOX mixing ratio needed to achieve this $f_{\text{obs,HOX}}$.

A1 Calculation of $f_{\text{obs,HOX}}$

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We assume all modeled sulfate in the MBL is formed via gas-phase OH oxidation and aqueous-phase H_2O_2 or O_3 oxidation based on the insignificant contribution (<3%) of other sulfate sources in the model.

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$$f_{\text{mod,OH}} + f_{\text{mod,H2O2}} + f_{\text{mod,O3}} = 1 \quad (\text{A1})$$

$$af_{\text{mod,H2O2}} + bf_{\text{mod,O3}} = \Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-}) \quad (\text{A2})$$

475 where $a = \Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{H2O2}} = 0.7 \text{ ‰}$ and $b = \Delta^{17}\text{O}(\text{nssSO}_4^{2-})_{\text{O3}} = 6.5 \text{ ‰}$. For the observations, we assume all sulfate in the MBL is formed via gas-phase OH oxidation, aqueous-phase H_2O_2 , O_3 and HOX oxidation pathways:

$$f_{\text{obs,OH}} + f_{\text{obs,H2O2}} + f_{\text{obs,O3}} + f_{\text{obs,HOX}} = 1 \quad (\text{A3})$$

$$af_{\text{obs,H2O2}} + bf_{\text{obs,O3}} = \Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-}) \quad (\text{A4})$$

480 where $f_{\text{obs,OH}}$, $f_{\text{obs,H2O2}}$, $f_{\text{obs,O3}}$ and $f_{\text{obs,HOX}}$ are the observed fractions of sulfate formed via gas-phase OH, aqueous-phase H_2O_2 , O_3 and HOX oxidation pathways, respectively. To solve for $f_{\text{obs,HOX}}$, two more equations involving $f_{\text{obs,OH}}$, $f_{\text{obs,H2O2}}$, $f_{\text{obs,O3}}$ and $f_{\text{obs,HOX}}$ are needed, in addition to Eqs. (A3-A4). Here we assume $f_{\text{obs,O3}}/f_{\text{obs,H2O2}}$ ratio is offset from $f_{\text{mod,O3}}/f_{\text{mod,H2O2}}$ ratio by Δr_1 and $f_{\text{obs,OH}}/f_{\text{obs,H2O2}}$ ratio is offset from $f_{\text{mod,OH}}/f_{\text{mod,H2O2}}$ ratio by Δr_2 :

$$\frac{f_{\text{obs,O3}}}{f_{\text{obs,H2O2}}} = \frac{f_{\text{mod,O3}}}{f_{\text{mod,H2O2}}} + \Delta r_1 \quad (\text{A5})$$

$$\frac{f_{\text{obs,OH}}}{f_{\text{obs,H2O2}}} = \frac{f_{\text{mod,OH}}}{f_{\text{mod,H2O2}}} + \Delta r_2 \quad (\text{A6})$$

485 Combining Eqs. (A3-A6) and using Eqs. (A1-A2) yield:

$$f_{\text{obs,HOX}} = 1 - \frac{\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})} + \Delta f \quad (\text{A7})$$

where

$$\Delta f = f_{\text{obs,H2O2}} \left[\left(\frac{b}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})} - 1 \right) \Delta r_1 - \Delta r_2 \right] \quad (\text{A8})$$

and

$$490 \quad f_{\text{obs,H2O2}} = \frac{f_{\text{mod,H2O2}} \Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-}) + b \Delta r_1 f_{\text{mod,H2O2}}} \quad (\text{A9})$$

Setting $\Delta f=0$ yields Eq. (8) in Sect. 5.2.4. Δf is zero when $\Delta r_1 = \Delta r_2 = 0$, which effectively assumes that the decreases in $f_{\text{mod,OH}}$, $f_{\text{mod,H2O2}}$ and $f_{\text{mod,O3}}$ after adding HOX in the model could be proportional to their relative fractions in the model. Δr_1 will be zero if cloud pH is unchanged, i.e. the S(IV) partitioning will remain unchanged after adding “S(IV) + HOX” reaction. The potential magnitude of Δf , which is dependent on the relative magnitude of Δr_1 and Δr_2 , is discussed below.

495 Δr_1 is expected to be negative with the addition of “S(IV) + HOX” reaction. Additional sulfate production in the aqueous phase will decrease cloud pH, resulting in decreases in the fractional contribution of O_3 relative to H_2O_2 ($f_{\text{O3}}/f_{\text{H2O2}}$). [The](#)

magnitude of the potential decrease in cloud pH can only be obtained after adding the “S(IV) + HOX” reactions in the model, which will be done in a follow-up study.

Δr_2 is expected to be positive with the addition of “S(IV) + HOX” reaction. HOX competes with H_2O_2 during oxidation of HSO_3^- in clouds, which causes a direct decrease in the fraction of sulfate formed via H_2O_2 oxidation ($f_{\text{mod,H}_2\text{O}_2}$). On the other hand, gas-phase sulfate production from oxidation of SO_2 by OH occurs mainly in the absence of clouds. Thus, adding “S(IV) + HOX” reaction causes an indirect decrease in the fraction of sulfate formed via OH oxidation ($f_{\text{mod,OH}}$) by increasing in-cloud and consequent total sulfate production, which depends on the availability of S(IV). Our model indicates that in-cloud sulfate production is limited by S(IV) abundance among our sampling locations (see the doubling H_2O_2 and O_3 simulation below), such that the decrease of $f_{\text{mod,OH}}$ is small compared to that of $f_{\text{mod,H}_2\text{O}_2}$ with the addition of “S(IV) + HOX” reaction, which results in an increase in the fractional contribution of OH relative to H_2O_2 ($f_{\text{OH}}/f_{\text{H}_2\text{O}_2}$).

As $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-}) < b$, the term $(\frac{b}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})} - 1)$ is positive. In addition, $f_{\text{obs,H}_2\text{O}_2}$ is positive. Thus, Δf is likely a negative number, which indicates Eq. (8) in Sect. 5.2.4 may overestimate $f_{\text{obs,HOX}}$.

We simulate the effect of an additional aqueous-phase reaction (“S(IV) + HOX”) in the model on Δr_1 and Δr_2 by doubling both H_2O_2 concentration and O_3 concentration. By doing this, we attribute half of the sulfate produced via H_2O_2 and O_3 oxidation to HOX oxidation. This simulation yields fractional contribution of sulfate formed via gas-phase OH, aqueous-phase H_2O_2 , O_3 and HOX oxidation pathways to be 0.17, 0.33, 0.08 and 0.40, respectively, among all sampling locations. Compared to the standard run, $f_{\text{mod,O}_3}/f_{\text{mod,H}_2\text{O}_2}$ ratio decreases from 0.35 to 0.23 and $f_{\text{mod,OH}}/f_{\text{mod,H}_2\text{O}_2}$ ratio increases from 0.35 to 0.53 on average among all sampling locations, yielding $\Delta r_1 = -0.12$ and $\Delta r_2 = 0.18$. Substituting standard run results ($\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-}) = 0.7 \text{‰}$, $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-}) = 1.8 \text{‰}$ and $f_{\text{mod,H}_2\text{O}_2} = 0.57$ (Table 2)) and $\Delta r_1 = -0.12$ and $\Delta r_2 = 0.18$ into Eq. (A8-A9) yields $\Delta f = -0.15$. Thus, if the addition of “S(IV) + HOX” reaction results in 40% of sulfate formed via oxidation of S(IV) by HOX, then our estimate of $f_{\text{obs,HOX}}$ using Eq. (8) in Sect. 5.4.2 would be 15% too high. The fraction (40%) of sulfate formed via oxidation of S(IV) by HOX is within the range of the averaged $f_{\text{obs,HOX}}$ (33-50%) calculated from observed and modeled $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ using Eq. (8). Thus, we suggest that calculating $f_{\text{obs,HOX}}$ for our samples using Eq. (8) may overestimate $f_{\text{obs,HOX}}$ by about 15%. This is smaller than the difference in $f_{\text{obs,HOX}}$ calculated from our low cloud pH (0.33 ± 0.32) and standard model (0.50 ± 0.33) runs. The actual magnitude of Δf can only be obtained by implementing the “S(IV)+HOX” reaction in the model.

A2 Calculation of HOX mixing ratios

We estimate the mixing ratio of HOX needed to achieve $f_{\text{obs,HOX}}$. First, we divide $f_{\text{obs,HOX}}$ into two parts:

$$f_{\text{obs,HOX}} = f_{1,\text{HOX}} + f_{2,\text{HOX}} \quad (\text{A10})$$

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530 where $f_{1,\text{HOX}}$ and $f_{2,\text{HOX}}$ are fractional contributions from “HOX + HSO₃⁻” reaction and “HOX + SO₃²⁻” reaction, respectively. H₂O₂ reacts with HSO₃⁻ only and O₃ reacts mainly with SO₃²⁻ (Hoffmann and Calvert, 1985), while HOX reacts quickly with both HSO₃⁻ and SO₃²⁻. Then we compare $f_{1,\text{HOX}}$ with $f_{\text{obs,H2O2}}$ and $f_{2,\text{HOX}}$ with $f_{\text{obs,O3}}$:

$$\frac{f_{1,\text{HOX}}}{f_{\text{obs,H2O2}}} = \frac{k_{\text{HOX+HSO}_3^-}[\text{HOX}]_{\text{aq}}}{k_{\text{H2O2+HSO}_3^-}[\text{H}_2\text{O}_2]_{\text{aq}}} \quad (\text{A11})$$

$$\frac{f_{2,\text{HOX}}}{f_{\text{obs,O3}}} = \frac{k_{\text{HOX+SO}_3^{2-}}[\text{HOX}]_{\text{aq}}}{k_{\text{O}_3+\text{SO}_3^{2-}}[\text{O}_3]_{\text{aq}}} \quad (\text{A12})$$

535 where $k_{\text{H2O2+HSO}_3^-}$ and $k_{\text{HOX+HSO}_3^-}$ are rate coefficients for reactions of H₂O₂ and HOX with HSO₃⁻, respectively; $k_{\text{O}_3+\text{SO}_3^{2-}}$ and $k_{\text{HOX+SO}_3^{2-}}$ are rate coefficients for reactions of O₃ and HOX with SO₃²⁻, respectively; and [H₂O₂]_{aq}, [O₃]_{aq} and [HOX]_{aq} are the aqueous-phase concentration of H₂O₂, O₃ and HOX in the cloud droplets, respectively. Combining Eqs. (A10-A12) yields:

$$[\text{HOX}]_{\text{aq}} = \frac{f_{\text{obs,HOX}}}{\frac{k_{\text{HOX+HSO}_3^-}f_{\text{obs,H2O2}}}{k_{\text{H2O2+HSO}_3^-}[\text{H}_2\text{O}_2]_{\text{aq}}} + \frac{k_{\text{HOX+SO}_3^{2-}}f_{\text{obs,O3}}}{k_{\text{O}_3+\text{SO}_3^{2-}}[\text{O}_3]_{\text{aq}}}} \quad (\text{A13})$$

This is the same equation as Eq. (9) in Sect. 5.4.2. $f_{\text{obs,HOX}}$ and $f_{\text{obs,H2O2}}$ are calculated in Eq. (A7) and Eq. (A9), respectively. Combining Eqs. (A5 and A9) yields:

$$540 \quad f_{\text{obs,O3}} = \frac{(f_{\text{mod,O3}} + \Delta r_1 f_{\text{mod,H2O2}}) \Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})}{\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-}) + 6.5 \Delta r_1 f_{\text{mod,H2O2}}} \quad (\text{A14})$$

For the calculations of $f_{\text{obs,HOX}}$, $f_{\text{obs,H2O2}}$ and $f_{\text{obs,O3}}$ in Sect. 5.4.2, both Δr_1 and Δr_2 are assumed to be zero.

Reference

- 545 Alexander, B., Park, R. J., Jacob, D. J., Li, Q. B., Yantosca, R. M., Savarino, J., Lee, C. C. W., and Thiemens, M. H.: Sulfate formation in sea salt aerosols: Constraints from oxygen isotopes, *J. Geophys. Res.*, 110, D10307, doi:10.1029/2004JD005659, 2005.
- Alexander, B., Park, R. J., Jacob, D. J., and Gong, S.: Transition metal-catalyzed oxidation of atmospheric sulfur: global implications for the sulfur budget, *J. Geophys. Res.*, 114, D02309, doi:10.1029/2008JD010486, 2009.
- 550 Alexander, B., Allman, D. J., Amos, H. M., Fairlie, T. D., Dachs, J., Hegg, D. A., and Sletten, R. S.: Isotopic constraints on the formation pathways of sulfate aerosol in the marine boundary layer of the subtropical northeast Atlantic Ocean, *J. Geophys. Res.*, 117, D06304, doi:10.1029/2011JD016773, 2012.

- Bao, H. M., Thiemens, M. H., Farquhar, J., Campbell, D. A., Lee, C. C. W., Heine, K., and Loope, D. B.: Anomalous ^{17}O compositions in massive sulphate deposits on the Earth, *Nature*, 406, 176-178, doi:10.1038/35018052, 2000.
- 555 Bates, T. S., Lamb, B. K., Guenther, A., Dignon, J., and Stoiber, R. E.: Sulfur emissions to the atmosphere from natural sources, *J. Atmos. Chem.*, 14, 315–337, doi:10.1007/BF00115242, 1992.
- Blatchley III, E. R., Johnson, R. W., Alleman, J. E., and McCoy, W. F.: Effective Henry's law constant for free chlorine and free bromine, *Wat. Res.*, 26, 99–106, doi:10.1016/0043-1354(92)90117-M, 1992.
- Barkan, E. and Luz, B.: High precision measurements of $^{17}\text{O}/^{16}\text{O}$ and $^{18}\text{O}/^{16}\text{O}$ ratios in H_2O , *Rapid Commun. Mass. Sp.*, 19, 3737–42, doi:10.1002/rem.2250, 2005.
- 560 Boucher, O., Moulin, C., Belviso, S., Aumont, O., Bopp, L., Cosme, E., vonKuhlmann, R., Lawrence, M. G., Pham, M., Reddy, M. S., Sciare, J., and Venkataraman, C.: DMS atmospheric concentrations and sulphate aerosol indirect radiative forcing: a sensitivity study to the DMS source representation and oxidation, *Atmos. Chem. Phys.*, 3, 49–65, doi:10.5194/acp-3-49-2003, 2003.
- 565 Carslaw, K. S., Lee, L. A., Reddington, C. L., Pringle, K. J., Rap, A., Forster, P. M., Mann, G. W., Spracklen, D. V., Woodhouse, M. T., Regayre, L. A., and Pierce, J. R.: Large contribution of natural aerosols to uncertainty in indirect forcing, *Nature*, 503, 67–71, doi:10.1038/nature12674, 2013.
- Chameides, W. L. and Stelson, A. W.: Aqueous-phase chemical processes in deliquescent sea-salt aerosols: A mechanism that couples the atmospheric cycles of S and sea-salt, *J. Geophys. Res.*, 97, 20565–20580, doi:10.1029/92JD01923, 1992.
- 570 Chin, M., Rood, R. B., Lin, S.-J., Muller, J. F., and Thompson, A. M.: Atmospheric sulfur cycle in the global model GOCART: Model description and global properties, *J. Geophys. Res.*, 105, 24 671–24 687, doi: 10.1029/2000JD900384, 2000.
- Cole-Dai, J., Thompson, L.G. and Mosley-Thompson, E.: A 485-year record of atmospheric chloride, nitrate and sulfate: Results of chemical analysis of ice cores from Dyer Plateau, Antarctic Peninsula, *Annals of Glaciology*, 21, 182-188, 1995.
- 575 Dominguez, G., Jackson, T., Brothers, L., Barnett, B., Nguyen, B., and Thiemens, M. H.: Discovery and measurement of an isotopically distinct source of sulfate in Earth's atmosphere, *P. Natl. Acad. Sci. USA*, 105(35), 12769–12773, doi:10.1073/pnas.0805255105, 2008.
- 580 Dubey, M.K., Mohrsladt, R., Donahue, N. M., and Anderson, J. G.: Isotope specific kinetics of hydroxyl radical (OH) with water (H_2O): Testing models of reactivity and atmospheric fractionation, *J. Phys. Chem. A.*, 101, 1494–1500, doi: 10.1021/jp962332p, 1997.

Fahey, K.M. and Pandis, S.N.: Size-resolved aqueous-phase atmospheric chemistry in a three-dimensional chemical transport model, *J. Geophys. Res.*, 108(D22), 4690, doi:10.1029/2003JD003564, 2003.

Faloona, I.: Sulfur processing in the marine atmospheric boundary layer: A review and critical assessment of modeling uncertainties, *Atmos. Environ.*, 43(18), 2841–2854, ISSN 1352-2310, doi:10.1016/j.atmosenv.2009.02.043, 2009.

585

[Faloona, I., Tan, D., Brune, W., Hurst, J., Barkot, D., Couch, T.L., Shepson, P., Apel, E., Riemer, D., Thornberry, T. and Carroll, M.A.: Nighttime observations of anomalously high levels of hydroxyl radicals above a deciduous forest canopy, *J. Geophys. Res.: Atmospheres*, 106\(D20\), 24315-24333, doi: 10.1029/2000JD900691, 2001.](#)

[Fickert, S., Adams, J. W. and Crowley, J. N.: Activation of Br₂ and BrCl via uptake of HOBr onto aqueous salt solutions, *J. Geophys. Res.*, 104\(D19\), 23719-23727, doi: 10.1029/1999JD900359, 1999.](#)

590

Fogelman, K. D., Walker, D. M., and Margerum, D. W.: Nonmetal redox kinetics: hypochlorite and hypochlorous acid reactions with sulfite, *Inorg. Chem.*, 28, 986–993, doi: 10.1021/ic00305a002, 1989.

Frenzel, A., Scheer, V., Sikorski, R., George, C., Behnke, W., and Zetzsch, C.: Heterogeneous Interconversion Reactions of BrNO₂, ClNO₂, Br₂, and Cl₂, *J. Phys. Chem. A*, 102, 1329–1337, doi: 10.1021/jp973044b, 1998.

595

Geng, L., Schauer, A.J., Kunasek, S.A., Sofen, E.D., Erbland, J., Savarino, J., Allman, D.J., Sletten, R.S., and Alexander, B.: Analysis of oxygen-17 excess of nitrate and sulfate at sub-micromole levels using the pyrolysis method, *Rapid Comm. Mass Spectrom.*, 27(21), 2411-2419, doi:10.1002/rcm.6703, 2013.

600

Gonfiantini, R., Stichler, W., and Rozanski, K.: Standards and intercomparison materials distributed by the International Atomic Energy Agency for stable isotope measurements, in: Reference and intercomparison materials for stable isotopes of light elements: Proceedings of a consultants meeting held in Vienna, 1–3 December 1993, IAEA-TECDOC-825, International Atomic Energy Agency, Vienna, 1993.

González-Gaya, B., Zúñiga-Rival, J., Ojeda, M. J., Jiménez, B. and Dachs, J.: Field measurements of the atmospheric dry deposition fluxes and velocities of polycyclic aromatic hydrocarbons to the global oceans, *Environ. Sci. Technol.*, 48, 5583-5592, doi: 10.1021/es503490z, 2014.

605

Hanson, D. R. and Ravishankara, A. R.: The reaction probabilities of ClONO₂ and N₂O₅ on 40 to 75% sulfuric acid solutions, *J. Geophys. Res.*, 96D, 17307–17314, doi: 10.1029/91JD01750, 1991.

Harris, E., Sinha, B., van Pinxteren, D., Tilgner, A., Fomba, W., Schneider, J., Roth, A., Gnauk, T., Fahlbusch, B., Mertes, S., Lee, T., Collett, J., Foley, S., Borrmann, S., Hoppe, P., and Herrmann, H.: Enhanced role of transition metal ion catalysis during in-cloud oxidation of SO₂, *Science*, 340, 727–730, doi:10.1126/science.1230911, 2013.

- 610 Harris, E., Sinha, B., van Pinxteren, D., Schneider, J., Poulain, L., Collett, J., D'Anna, B., Fahlbusch, B., Foley, S., Fomba, K. W., George, C., Gnauk, T., Henning, S., Lee, T., Mertes, S., Roth, A., Stratmann, F., Borrmann, S., Hoppe, P., and Herrmann, H.: In-cloud sulfate addition to single particles resolved with sulfur isotope analysis during HCCT-2010, *Atmos. Chem. Phys.*, 14, 4219–4235, doi:10.5194/acp-14-4219-2014, 2014.
- Haywood, J. and Boucher, O.: Estimates of the direct and indirect radiative forcing due to tropospheric aerosols: a review, *Rev. Geophys.*, 38, 513–543, doi: 10.1029/1999RG000078, 2000.
- 615 Hegg, D. A., Yuen, P.-F. and Larson, T. V.: Modeling the effects of heterogeneous cloud chemistry on the marine particle size distribution, *J. Geophys. Res.*, 97(D12), 12927–12933, doi: 10.1029/92JD01184, 1992.
- Holzwarth, G., Balmer, R. G., and Soni, L.: The fate of chlorine and chloramines in cooling towers Henry's law constants for flashoff, *Wat. Res.*, 18, 1421–1427, doi: 10.1016/0043-1354(84)90012-5, 1984.
- 620 Huthwelker, T., Clegg, S. L., Peter, T., Carslaw, K., Luo, B. P., and Brimblecombe, P.: Solubility of HOCl in water and aqueous H₂SO₄ to stratospheric temperatures, *J. Atmos. Chem.*, 21, 81–95, doi: 10.1007/BF00712439, 1995.
- Jauhainen, J., Moore, J., Perämäki, P., Derome, J., Derome, K.: Simple procedure for ion chromatographic determination of anions and cations at trace levels in ice core samples, *Analytical Chimica Acta*, 389, 21–29, doi:10.1016/S0003-2670(99)00185-3, 1999.
- Jenkins, K. A. and Bao, H.: Multiple oxygen and sulfur isotope compositions of atmospheric sulfate in Baton Rouge, LA, USA, *Atmos. Environ.*, 40, 4528–4537, doi:10.1016/j.atmosenv.2006.04.010, 2006.
- 625 Kerminen, V.-M., Petäjä, T., Manninen, H. E., Paasonen, P., Nieminen, T., Sipilä, M., Junninen, H., Ehn, M., Gagné, S., Laakso, L., Riipinen, I., Vehkamäki, H., Kurten, T., Ortega, I. K., Dal Maso, M., Brus, D., Hyvärinen, A., Lihavainen, H., Leppä, J., Lehtinen, K. E. J., Mirme, A., Mirme, S., Hörrak, U., Berndt, T., Stratmann, F., Birmili, W., Wiedensohler, A., Metzger, A., Dommen, J., Baltensperger, U., Kiendler-Scharr, A., Mentel, T. F., Wildt, J., Winkler, P. M., Wagner, P. E.,
- 630 Petzold, A., Minikin, A., Plass-Dülmer, C., Pöschl, U., Laaksonen, A., and Kulmala, M.: Atmospheric nucleation: highlights of the EUCAARI project and future directions, *Atmos. Chem. Phys.*, 10, 10829–10848, doi:10.5194/acp-10-10829-2010, 2010.
- Kettle, A., Andreae, M., Amouroux, D., Andreae, T., Bates, T., Berresheim, H., Bingemer, H., Boniforti, R., Curran, M., DiTullio, G., Helas, G., Jones, G., Keller, M., Kiene, R., Leck, C., Lévasseur, M., Malin, G., Maspero, M., Matrai, P., McTaggart, A., Mihalopoulos, N., Nguyen, B., Novo, A., Putaud, J., Rapsomanikis, S., Roberts, G., Schebeske, G., Sharma, S., Simo, R., Staubes, R., Turner, S., and Uher, G.: A global database of sea surface dimethylsulfide (DMS) measurements and a procedure to predict sea surface DMS as a function of latitude, longitude, and month, *Global Biogeochem. Cy.*, 13, 399–444, doi: 10.1029/1999GB900004, 1999.

- 640 Kulmala, M., Pirjola, L. and Mäkelä, J.M.: Stable sulphate clusters as a source of new atmospheric particles, *Nature*, 404(6773), 66–69, doi:10.1038/35003550, 2000.
- Lawler, M. J., Sander, R., Carpenter, L. J., Lee, J. D., von Glasow, R., Sommariva, R., and Saltzman, E. S.: HOCl and Cl₂ observations in marine air, *Atmos. Chem. Phys.*, 11, 7617–7628, doi:10.5194/acp-11-7617-2011, 2011.
- Le Breton, M., Bannan, T., Shallcross, D. E., Khan, M. A. H., Evans, M. J., Lee, J., Lidster, R. T., Andrews, S. J., Carpenter, L. J., Sherwen, T., Schmidt, J. A., Jacob, D. J., Harris, N. R. P., Bauguitte, S. J., Gallagher, M. W., Bacak, A., Leather, K. E. and Percival C. J.: Enhanced ozone loss by active inorganic bromine chemistry in the tropical troposphere, in review, 2016.
- 645 Lee, C. C.-W. and Thiemens, M. H.: The $\delta^{17}\text{O}$ and $\delta^{18}\text{O}$ measurements of atmospheric sulfate from a coastal and high alpine region: A mass-independent isotopic anomaly, *J. Geophys. Res.*, 106(D15), 17359–17373, doi:10.1029/2000JD900805, 2001.
- 650 Lee, C. C. W., Savarino, J. H., Cachier, H. and Thiemens, M. H.: Sulfur (^{32}S , ^{33}S , ^{34}S , ^{36}S) and oxygen (^{16}O , ^{17}O , ^{18}O) isotopic ratios of primary sulfate produced from combustion processes, *Tellus B*, 54(3), 193-200, doi: 10.1034/j.16000889.2002.01384.x, 2002.
- Lelieveld, J. and Heintzenberg, J.: Sulfate cooling effect on climate through in-cloud oxidation of anthropogenic SO₂, *Science*, 258, 117–120, doi: 10.1126/science.258.5079.117, 1992.
- 655 Liao, J., Huey, L. G., Tanner, D. J., Flocke, F. M., Orlando, J. J., Neuman, J. A., Nowak, J. B., Weinheimer, A. J., Hall, S. R., Smith, J. N., Fried, A., Staebler, R. M., Wang, Y., Koo, J.-H., Cantrell, C. A., Weibring, P., Walega, J., Knapp, D. J., Shepson, P. B., and Stephens, C. R.: Observations of inorganic bromine (HOBr, BrO, and Br₂) speciation at Barrow, Alaska, in spring 2009, *J. Geophys. Res.*, 117, D00R16, doi:10.1029/2011JD016641, 2012.
- Liu, Q.: Kinetics of aqueous phase reactions related to ozone depletion in the arctic troposphere: bromine chloride hydrolysis, bromide ion with ozone, and sulfur(IV) with bromine and hypobromous acid, Ph.D. thesis, Department of Chemistry, Purdue University, USA, 253 pp., 2000.
- 660 [Lu, K. D., Rohrer, F., Holland, F., Fuchs, H., Brauers, T., Oebel, A., Dlugi, R., Hu, M., Li, X., Lou, S. R., Shao, M., Zhu, T., Wahner, A., Zhang, Y. H., and Hofzumahaus, A.: Nighttime observation and chemistry of HO_x in the Pearl River Delta and Beijing in summer 2006, *Atmos. Chem. Phys.*, 14, 4979-4999, doi:10.5194/acp-14-4979-2014, 2014.](#)
- 665 Lyons, J. R.: Transfer of mass-independent fractionation on ozone to other oxygen-containing molecules in the atmosphere, *Geophys. Res. Lett.*, 28, 3231–3234, doi: 10.1029/2000GL012791, 2001.

a 2016-7-19 5:04 PM

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- 670 Mauldin III, R. L., Berndt, T., Sipilä, M., Paasonen, P., Petäjä, T., Kim, S., Kurtén, T., Stratmann, F., Kerminen, V.-M., and Kulmala, M.: A new atmospherically relevant oxidant of sulphur dioxide, *Nature*, 488, 193–196, doi:10.1038/nature11278, 2012.
- McCabe, J. R., Savarino, J., Alexander, B., Gong, S. and Thiemens, M. H.: Isotopic constraints on non-photochemical sulfate production in the Arctic winter, *Geophys. Res. Lett.*, 33, L05810, doi:10.1029/2005GL025164, 2006.
- McCoy, W.F., Blatchley III E.R. and Johnson, R.W.: Hypohalous acid and haloamine flash off in industrial evaporative cooling systems, Technical paper number TP90-09, Cooling Tower Institute 1990 annual meeting, Cooling Technology Institute, Houston, TX, 1990.
- 675 Millero, F. J., Feistel, R., Wright, D. G., and McDougall, T. J.: The composition of Standard Seawater and the definition of the Reference-Composition Salinity Scale, *Deep-Sea Res. Pt. I*, 55, 50–72, doi:10.1016/j.dsr.2007.10.001, 2008.
- Myhre, G., Shindell, D., Bréon, F.-M., Collins, W., Fuglestedt, J., Huang, J., Koch, D., Lamarque, J.-F., Lee, D., Mendoza, B., Nakajima, T., Robock, A., Stephens, G., Takemura, T. and Zhang, H.: Anthropogenic and Natural Radiative Forcing. In: *Climate Change 2013: The Physical Science Basis. Contribution of Working Group I to the Fifth Assessment Report of the Intergovernmental Panel on Climate Change* [Stocker, T.F., D. Qin, G.-K. Plattner, M. Tignor, S.K. Allen, J. Boschung, A. Nauels, Y. Xia, V. Bex and P.M. Midgley (eds.)], Cambridge University Press, Cambridge, United Kingdom and New York, NY, USA, 2013.
- 680 Neuman, J. A., Nowak, J. B., Huey, L. G., Burkholder, J. B., Dibb, J. E., Holloway, J. S., Liao, J., Peischl, J., Roberts, J. M., Ryerson, T. B., Scheuer, E., Stark, H., Sticker, R. E., Tanner, D. J., and Weinheimer, A.: Bromine measurements in ozone depleted air over the Arctic Ocean, *Atmos. Chem. Phys.*, 10, 6503–6514, doi:10.5194/acp-10-6503-2010, 2010.
- 685 O’Dowd, C. D., Lowe, J. A., Clegg, N., Smith, M. H. and Clegg, S. L.: Modeling heterogeneous sulphate formation in maritime stratiform clouds, *J. Geophys. Res.*, 105(D6), 7143–7160, doi:10.1029/1999JD900915, 2000.
- Olivier, J. G. J. and Berdowski, J. J. M.: Global emissions sources and sinks. In: Berdowski, J., Guicherit, R. and B.J. Heij (eds.) “*The Climate System*”, pp. 33-78. A.A. Balkema Publishers/Swets & Zeitlinger Publishers, Lisse, The Netherlands, ISBN 90 5809 255 0, 2001.
- 690 Park, R. J., Jacob, D. J., Field, B. D., Yantosca, R. M., and Chin, M.: Natural and transboundary pollution influences on sulfate-nitrate-ammonium aerosols in the United States: implications for policy, *J. Geophys. Res.*, 109, D15204, doi:10.1029/2003JD004473, 2004.

- 695 Patris, N., Cliff, S. S., Quinn, P. K., Kasem, M., and Thiemens, M. H.: Isotopic analysis of aerosol sulfate and nitrate during ITCT-2k2: Determination of different formation pathways as a function of particle size, *J. Geophys. Res.-Atmos.*, 112, D23301, doi:10.1029/2005JD006214, 2007.
- Pierce, J. R., Evans, M. J., Scott, C. E., D'Andrea, S. D., Farmer, D. K., Swietlicki, E., and Spracklen, D. V.: Weak global sensitivity of cloud condensation nuclei and the aerosol indirect effect to Criegee + SO₂ chemistry, *Atmos. Chem. Phys.*, 13, 3163–3176, doi:10.5194/acp-13-3163-2013, 2013.
- 700 Roelofs, G. J. H.: A cloud chemistry sensitivity study and comparison of explicit and bulk cloud model performance, *Atmos. Environ., Part A*, 27(15), 2255–2264, doi: 10.1016/0960-1686(93)90396-G, 1993.
- Sander, R.: Compilation of Henry's law constants (version 4.0) for water as solvent, *Atmos. Chem. Phys.*, 15, 4399–4981, doi:10.5194/acp-15-4399-2015, 2015.
- 705 Sander, S. P., Friedl, R. R., Golden, D. M., Kurylo, M. J., Moortgat, G. K., Keller Rudek, H., Wine, P. H., Ravishankara, A. R., Kolb, C. E., Molina, M. J., Finlayson-Pitts, B. J., Huie, R. E., and Orkin, V. L.: Chemical Kinetics and Photochemical Data for Use in Atmospheric Studies, Evaluation Number 15, JPL Publication 06-2, Jet Propulsion Laboratory, Pasadena, CA, available at: <http://jpldataeval.jpl.nasa.gov> (last access: 25 March 2016), 2006
- Savarino, J. and Thiemens, M. H.: Analytical procedure to determine both $\delta^{18}\text{O}$ and $\delta^{17}\text{O}$ of H₂O₂ in natural water and first measurements, *Atmos. Environ.*, 33, 3683–3690, doi:10.1016/S1352-2310(99)00122-3, 1999.
- 710 Savarino, J., Lee, C.C.W. and Thiemens M.H.: Laboratory oxygen isotopic study of sulfur (IV) oxidation: Origin of the mass-independent isotopic anomaly in atmospheric sulfates and sulfate mineral deposits on Earth, *J. Geophys. Res.*, 105, 29,079–29,088, doi:10.1029/ 2000JD900456, 2000.
- Schmidt, J. A., Jacob, D. J., Horowitz, H. M., Hu, L., Sherwen, T., Evans, M. J., Liang, Q., Suleiman, R. M., Oram, D. E., Le Breton, M., Percival, C. J., Wang, S., Dix, B., and Volkamer, R.: Modeling the observed tropospheric BrO background: Importance of multiphase chemistry and implications for ozone, OH, and mercury, *J. Geophys. Res.-Atmos.*, in press, 2016.
- 715 Sievering, H., Boatman, J., Galloway, J., Keene, W., Kim, Y., Luria, M. and Ray, J.: Heterogeneous sulfur conversion in sea-salt aerosol particles: the role of aerosol water content and size distribution, *Atmos. Environ. Part A. General Topics*, 25(8), 1479-1487, doi:10.1016/0960-1686(91)90007-T, 1991.
- 720 Sievering, H., Caine, J., Harvey, M., McGregor, J., Nichol, S., and Quinn, P.: Aerosol non-sea-salt sulphate in the remote marine boundary layer under clear-sky and normal cloudiness conditions: Ocean-derived biogenic alkalinity enhances sea-salt sulfate production by ozone oxidation, *J. Geophys. Res.*, 109, D19317, doi:10.1029/2003JD004315, 2004.

- Sofen, E. D., Alexander, B., and Kunasek, S. A.: The impact of anthropogenic emissions on atmospheric sulfate production pathways, oxidants, and ice core $\Delta^{17}\text{O}(\text{SO}_4^{2-})$, *Atmos. Chem. Phys.*, 11, 3565-3578, doi:10.5194/acp-11-3565-2011, 2011.
- 725 Stockwell, W. R. and Calvert, J. G.: The mechanism of the HO-SO₂ reaction, *Atmos. Environ.*, 17, 2231–2235, doi:10.1016/0004-6981(83)90220-2, 1983.
- Streets, D. G., Zhang, Q., Wang, L., He, K., Hao, J., Wu, Y., Tang, Y., and Carmichael, G. R.: Revisiting China's CO emissions after the Transport and Chemical Evolution over the Pa-cific (TRACE-P) mission: Synthesis of inventories, atmospheric modeling, and observations, *J. Geophys. Res.*, 111, D14306, doi:10.1029/2006JD007118, 2006.
- 730 Troy, R. C. and Margerum, D. W.: Non-metal redox kinetics: Hypobromite and hypobromous acid reactions with iodide and with sulfite and the hydrolysis of bromosulfate, *Inorg. Chem.*, 30, 3538–3543, doi:10.1021/ic00018a028, 1991.
- Vestreng, V. and Klein, H.: Emission data reported to UN- ECE/EMEP: Quality insurance and trend analysis & Presentation of WebDab, MSC-W Status Report, Tech. rep., Norwegian Meteorological Institute, Oslo, Norway, http://www.emep.int/publ/reports/2002/mscw_note_1_2002.pdf, 2002.
- 735 Vicars, W. C. and Savarino, J.: Quantitative constraints on the ¹⁷O-excess ($\Delta^{17}\text{O}$) signature of surface ozone: Ambient measurements from 50°N to 50°S using the nitrite-coated filter technique, *Geochim. Cosmochim. Acta*, 135, 270–287, doi:10.1016/j.gca.2014.03.023, 2014.
- Vogt, R., Crutzen, P. J., and Sander, R.: A mechanism for halogen release from sea-salt aerosol in the remote marine boundary layer, *Nature*, 383, 327–330, doi:10.1038/383327a0, 1996.
- 740 van Donkelaar, A., Martin, R. V., Pasch, A. N., Szykman, J. J., Zhang, L., Wang, Y. X., and Chen, D.: Improving the accuracy of daily satellite-derived ground-level fine aerosol concentration estimates for North America, *Environ. Sci. Technol.*, 46, 11971–11978, doi:10.1021/es3025319, 2012
- von Glasow, R., Sander, R., Bott, A., and Crutzen, P. J.: Modeling halogen chemistry in the marine boundary layer. 2. Interactions with sulfur and the cloud-covered MBL, *J. Geophys. Res.*, 107D, 4323, 10.1029/2001JD000943, 2002.
- 745 Yiin, B. S. and Margerum, D. W.: Kinetics of Hydrolysis of the Chlorosulfate Ion, *Inorg. Chem.*, 27, 1670–1672, doi:10.1021/ic00283a002, 1988.
- Yuen, P.-F., Hegg, D.A., Larson, T.V. and Barth, M.C.: Parameterization of heterogeneous droplet chemistry for use in bulk cloud models, *J. Appl. Meteorol.*, 35(5), 679–689, doi:10.1175/1520-0450(1996)035<0679:POHDCF>2.0.CO;2, 1996.

Tables

Table 1. $\Delta^{17}\text{O}$ of sulfate produced via different pathways.

Sulfate formation pathway	$\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ (‰)	References
$\text{SO}_2(\text{g})+\text{OH}$	0	Dubey et al., 1997; Lyons, 2001
$\text{S(IV)}+\text{H}_2\text{O}_2$	0.7	Savarino and Thiemens, 1999
$\text{S(IV)}+\text{O}_3$	6.5	Vicars and Savarino, 2014
$\text{S(IV)}+\text{O}_2$	-0.09	Barkan and Luz, 2005
$\text{S(IV)}+\text{HOX}$	0	Fogelman et al., 1989; Troy and Margerum, 1991

Table 2. $[\text{SO}_4^{2-}]_{\text{NSS}}$, $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$, $f_{\text{O}_3,\text{min}}$ and $f_{\text{O}_3,\text{max}}$ obtained from observations and $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$, $f_{\text{mod},\text{O}_3}$, $f_{\text{mod},\text{H}_2\text{O}_2}$ and $f_{\text{mod},\text{OH}}$ obtained from the standard model run for samples in different categories.

Category	Characteristic	Number	$[\text{SO}_4^{2-}]_{\text{NSS}}$ ($\mu\text{g m}^{-3}$)	$\Delta^{17}\text{O}_{\text{obs}}$ (‰)	$f_{\text{O}_3,\text{min}}$	$f_{\text{O}_3,\text{max}}$	$\Delta^{17}\text{O}_{\text{mod}}$ (‰)	$f_{\text{mod},\text{OH}}$	$f_{\text{mod},\text{H}_2\text{O}_2}$	$f_{\text{mod},\text{O}_3}$	$f_{\text{mod},\text{het}}$
I	Southern Ocean	13	1.4±0.8	0.5±0.3	0.01±0.02	0.08±0.05	3.6±1.1	0.04±0.04	0.46±0.16	0.48±0.19	0.02±0.02
II	Antarctic coast	18	1.4±0.7	0.7±0.4	0.03±0.04	0.11±0.06	1.1±0.3	0.25±0.11	0.65±0.09	0.10±0.04	0.01±0.01
III	Subtropical MBL	19	0.9±0.5	0.8±0.4	0.04±0.05	0.13±0.07	1.9±1.0	0.17±0.08	0.61±0.14	0.19±0.15	0.03±0.03
IV	Tropical coasts	17	1.3±0.9	0.8±0.4	0.03±0.04	0.13±0.06	1.2±0.4	0.30±0.16	0.54±0.15	0.10±0.06	0.03±0.03
Total	/	67	1.2±0.8	0.7±0.4	0.03±0.04	0.12±0.06	1.8±1.2	0.20±0.14	0.57±0.15	0.20±0.19	0.02±0.03

Table 3. $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$, $f_{\text{mod},\text{O}_3}$, $f_{\text{mod},\text{H}_2\text{O}_2}$ and $f_{\text{mod},\text{OH}}$ obtained from five sensitivity studies: (a) double OH concentration, (b) halve O_3 concentration, (c) double H_2O_2 concentration, (d) halve clouds and (e) low cloud pH.

Category	$\Delta^{17}\text{O}_{\text{mod}}$ (‰)	$f_{\text{mod},\text{OH}}$	$f_{\text{mod},\text{H}_2\text{O}_2}$	$f_{\text{mod},\text{O}_3}$	$f_{\text{mod},\text{het}}$
(a) 2[OH]					
I	3.8±1.1	0.06±0.05	0.39±0.14	0.53±0.18	0.02±0.02
II	1.3±0.4	0.29±0.11	0.57±0.09	0.13±0.05	0.01±0.01
III	1.8±1.0	0.25±0.11	0.51±0.13	0.20±0.16	0.03±0.03
IV	1.1±0.5	0.42±0.19	0.43±0.17	0.09±0.06	0.02±0.03
Total	1.9±1.3	0.27±0.17	0.48±0.15	0.22±0.20	0.02±0.02
(b) 1/2[O ₃]					
I	3.5±1.2	0.04±0.04	0.46±0.17	0.47±0.20	0.02±0.02
II	1.2±0.3	0.25±0.11	0.65±0.09	0.10±0.04	0.01±0.01
III	1.7±1.0	0.17±0.08	0.63±0.14	0.17±0.16	0.03±0.03
IV	1.1±0.4	0.31±0.16	0.56±0.16	0.08±0.05	0.03±0.03
Total	1.8±1.2	0.20±0.14	0.58±0.15	0.19±0.19	0.02±0.03
(c) 2[H ₂ O ₂]					
I	2.7±1.0	0.03±0.03	0.60±0.16	0.35±0.18	0.02±0.02
II	0.9±0.2	0.21±0.10	0.72±0.09	0.06±0.03	0.01±0.01
III	1.4±0.8	0.15±0.07	0.70±0.13	0.12±0.14	0.03±0.03
IV	1.0±0.3	0.27±0.15	0.61±0.16	0.07±0.03	0.03±0.03
Total	1.4±0.9	0.17±0.13	0.66±0.14	0.14±0.15	0.02±0.03
(d) 1/2 clouds					
I	3.7±1.1	0.06±0.05	0.42±0.14	0.50±0.18	0.02±0.03
II	1.1±0.3	0.31±0.12	0.58±0.11	0.10±0.04	0.01±0.01
III	1.9±1.0	0.23±0.10	0.53±0.13	0.19±0.15	0.04±0.04
IV	1.1±0.5	0.39±0.18	0.45±0.16	0.09±0.07	0.03±0.03
Total	1.8±1.2	0.26±0.17	0.50±0.15	0.20±0.19	0.03±0.03
(e) Low cloud pH					
I	1.3±0.4	0.07±0.03	0.81±0.09	0.08±0.06	0.03±0.04
II	0.6±0.1	0.26±0.10	0.72±0.10	0.01±0.00	0.01±0.01
III	1.4±0.5	0.18±0.07	0.68±0.10	0.10±0.05	0.03±0.04
IV	1.1±0.4	0.30±0.16	0.57±0.17	0.08±0.06	0.03±0.03
Total	1.1±0.5	0.21±0.13	0.69±0.14	0.07±0.06	0.02±0.03

Table 4. $f_{\text{obs,HOX}}$, $[\text{HOX}]_{\text{aq}}$ and $[\text{HOX}]_{\text{g}}$ calculated using the model results from (a) the standard run and (b) low cloud pH run, respectively. The mean and standard deviation are shown for $f_{\text{obs,HOX}}$. The range and median value are shown for $[\text{HOX}]_{\text{aq}}$ and $[\text{HOX}]_{\text{g}}$.

Category	(a) standard run			(b) low cloud pH		
	$f_{\text{obs,HOX}}$	$[\text{HOX}]_{\text{aq}}$ (pM)	$[\text{HOX}]_{\text{g}}$ (ppt)	$f_{\text{obs,HOX}}$	$[\text{HOX}]_{\text{aq}}$ (pM)	$[\text{HOX}]_{\text{g}}$ (ppt)
I	0.84±0.15	41~30192 (102)	0.04~30.97 (0.10)	0.58±0.30	1~4048 (27)	0.00~4.15 (0.03)
II	0.42±0.29	0~395 (26)	0.00~0.40 (0.03)	0.18±0.26	0~145 (0)	0.00~0.15 (0.00)
III	0.47±0.34	0~3334 (97)	0.00~3.42 (0.10)	0.36±0.32	0~2585 (40)	0.00~2.65 (0.04)
IV	0.35±0.29	0~761 (85)	0.00~0.78 (0.09)	0.28±0.31	0~764 (54)	0.00~0.78 (0.06)
Total	0.50±0.33	0~30192 (88)	0.00~30.97 (0.09)	0.33±0.32	0~4048 (11)	0.00~4.15 (0.01)

Figures

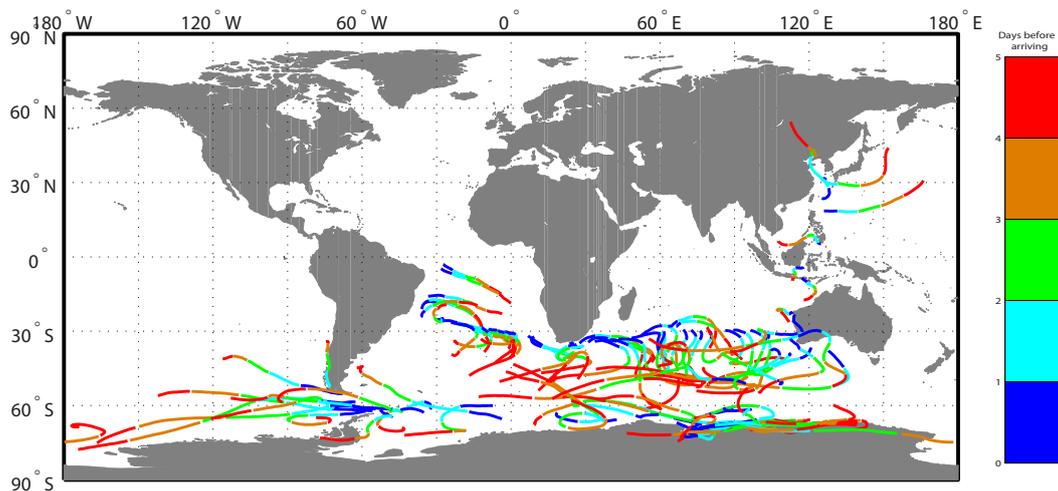


Figure 1. 5-day back trajectories calculated from the HYSPLIT model for all sampling locations. Blue indicates the ending point of each back trajectory.

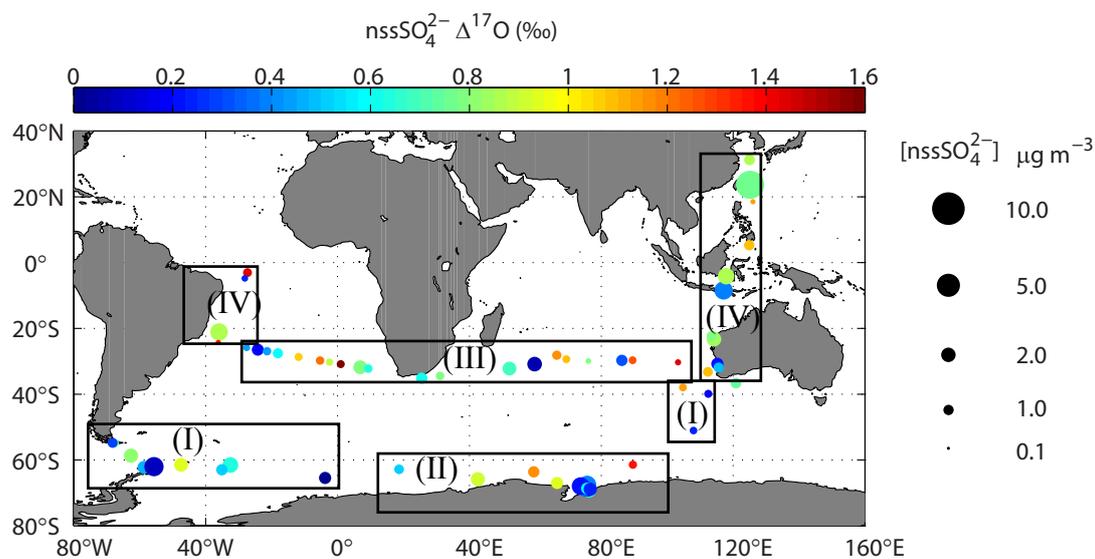


Figure 2. Observations of $\Delta^{17}\text{O}(\text{nssSO}_4^{2-})$ (‰) and nssSO_4^{2-} concentration ($\mu\text{g m}^{-3}$) for aerosol samples collected in the MBL during spring and summer in 2011 and 2012. Black rectangles indicate regions I – IV.

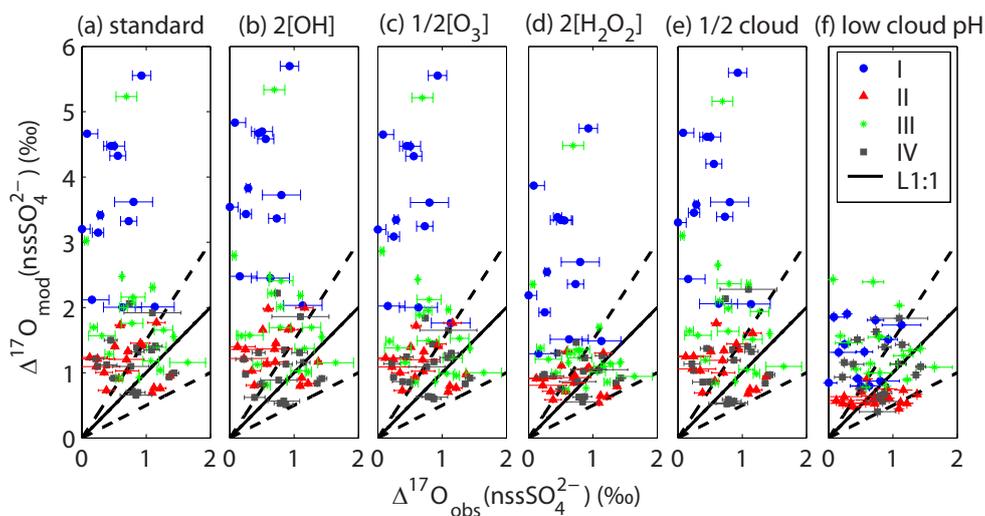


Figure 3. $\Delta^{17}\text{O}_{\text{mod}}(\text{nssSO}_4^{2-})$ versus $\Delta^{17}\text{O}_{\text{obs}}(\text{nssSO}_4^{2-})$ for different model simulations: (a) standard run, (b) double OH concentration, (c) halve O₃ concentration, (d) double H₂O₂ concentration, (e) halve clouds and (f) low cloud pH.

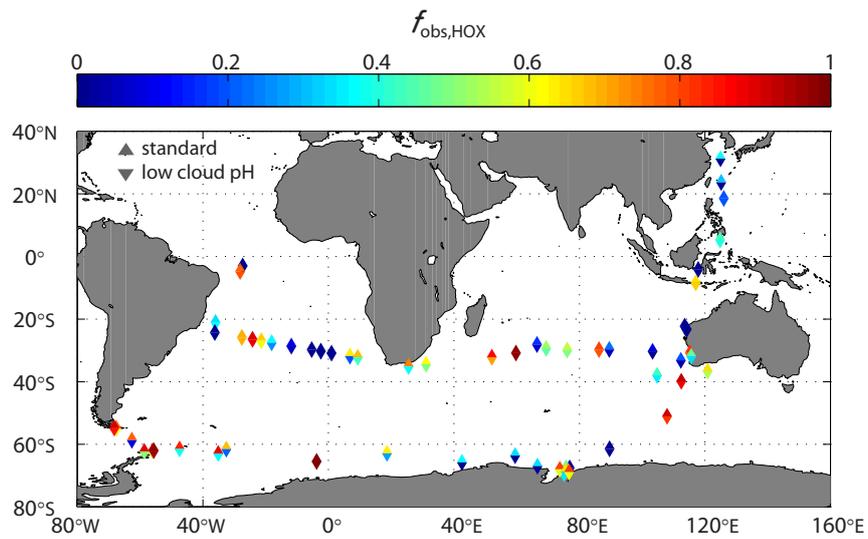


Figure 4. Calculated $f_{\text{obs,HOX}}$ for each sample using modeling results from the standard run (upward-pointing triangle) and low cloud pH run (downward-pointing triangle), respectively.