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SIZE DEPENDENT CHEMICAL AGEING OF OLEIC ACID AEROSOL UNDER DRY AND HUMIDIFIED CONDITIONS

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1 **ABSTRACT**

2 A chemical reaction chamber system has been developed for the processing of oleic acid aerosol

3 particles with ozone under two relative humidity conditions: dry and humidified to 65% R.H. The

4 apparatus consists of an aerosol flow tube, in which the ozonolysis occurs, coupled to a scanning

5 mobility particle sizer (SMPS) and an aerosol time-of-flight mass spectrometer (ATOFMS) which

6 measure the evolving particle size and composition. Under both relative humidity conditions,

7 ozonolysis results in a significant decrease in particle size and mass which is consistent with the

8 formation of volatile products that partition from the particle to the gas phase. Mass spectra derived

9 from the ATOFMS reveal the presence of the typically observed reaction products: azaleic acid,

10 nonanal, oxononanoic acid and nonanoic acid, as well as a range of higher molecular weight

11 products deriving from the reactions of reaction intermediates with oleic acid and its oxidation

12 products. These include octanoic acid, and 9- and 10-oxooctadecanoic acid, as well as products of

13 considerably higher molecular weight. Quantitative evaluation of product yields with the ATOFMS

14 shows a marked dependence upon both particle size association (from 0.3 to 2.1 µm diameter) and

15 relative humidity. Under both relative humidity conditions, the percentage residual of oleic acid

16 increases with increasing particle size and the main lower molecular weight products are nonanal

17 and oxonononic acid. Under dry conditions, the percentage of higher molecular weight products

18 increases with increasing particle size, due to the poorer internal mixing of the larger particles.

19 Under humidified conditions, the percentage of unreacted oleic acid is greater, except in the

20 smallest particle fraction, with little formation of high molecular weight products relative to the dry

21 particles. It is postulated that water reacts with reactive intermediates, competing with the

22 processes which produce high molecular weight products. Whilst the oleic acid model aerosol

23 system is of limited relevance to complex internally mixed atmospheric aerosol, the generic

24 findings presented in this paper give useful insights into the nature of heterogeneous chemical

25 processes.

26 **Keywords:** Heterogeneous reactions; oleic acid; ozone; ATOFMS

1.   INTRODUCTION

Atmospheric aerosol particles play critical roles in air quality, visibility, human health, regional and global climate, cloud condensation nuclei ability, precipitation events, atmospheric acid deposition, optical properties, atmospheric energy balance, and stratospheric ozone depletion (Harrison, 2014). Aerosol particles are typically composed of a mixture of inorganic and organic material. The organic component of aerosol is highly complex and may contain thousands of different chemical species of both biogenic and anthropogenic origin (Goldstein and Galbally, 2007). These organic components cause aerosol particles to exhibit a wide range of chemical properties due to their differing composition.

Primary organic aerosol (POA) particles are emitted directly from anthropogenic and biological sources (Pöschl, 2011). It is composed of a wide range of hydrocarbons, partially oxidized organics, and elemental carbon primarily related to combustion processes including burning of fossil fuels, cooking, domestic heating, and biomass burning. Natural biogenic sources of organic aerosol particles include plants and vegetation, the ocean surface, volcanic eruptions, and wind-driven dust (Pöschl, 2011). By contrast, secondary organic aerosol (SOA) particles are formed in the atmosphere from biogenic and anthropogenic gaseous precursors. Several pathways for the transformation of aerosol particles in the atmosphere have been identified, which may alter their chemical and physical properties, in general causing hydrophobic-to-hydrophilic conversion of organic components. These pathways include gas-phase reactions, condensed-phase reactions, multiphase reactions, and multigenerational chemistry which has received attention recently because it is the least understood (Rudich, 2003; Petters et al., 2006; Kroll and Seinfeld, 2008; Carlton et al., 2010; Koop et al., 2011; Kolb and Worsnop, 2012).

The ozonolysis of oleic acid (OL) aerosol is a much studied heterogeneous reaction which provides a readily accessible test system for the understanding of atmospheric processing of organic aerosol

1   under a range of environmental conditions (Hearn and Smith, 2004; Katrib et al., 2004; Ziemann,

2   2005; Gonzalez-Labrada et al., 2007; Zahardis and Petrucci, 2007; Vesna et al., 2009; Pfrang et al.,

3   2010; Lee et al., 2012; Chan et al., 2013; Hosny et al., 2013; Mendez et al., 2014; Hosny et al.,

4   2016) - although oleic acid itself is only introduced into the atmosphere in small quantities via the

5   heating of fat and cooking oil. OL studies have, for example, explored the effects of the droplet

6   state (Katrib et al., 2005a), relative humidity (RH) (Vesna et al., 2009; Lee et al., 2012), OL and

7   ozone concentration (Lee and Chan, 2007; Mendez et al., 2014), and extent of chemical ageing

8   (Reynolds et al., 2006) upon the OL - ozone system. The kinetics (Moise and Rudich, 2002;

9   Gonzalez-Labrada et al., 2007), reaction mechanism and products (Hearn and Smith, 2004; Katrib

10 et al., 2004; Hung et al., 2005; Zahardis et al., 2005; Ziemann, 2005; Zahardis et al., 2006; Zahardis

11 and Petrucci, 2007, Wang et al., 2016), particle morphology and hygroscopicity (Dennis-Smither et

12 al., 2012a), and viscosity (Hosny et al., 2013; 2016) have also been investigated.

13

14 The initial stages in the mechanism of particulate phase OL ozonolysis are comparatively well

15 understood (Moise and Rudich, 2002; Smith et al., 2002; Hearn and Smith, 2004; Katrib et al.,

16 2004; Ziemann, 2005; Grimm et al., 2006; Nash et al., 2006; Hung and Ariya, 2007; Zahardis and

17 Petrucci, 2007; Vesna et al., 2008; Lee et al., 2012), see Scheme 1. The initial step in the reaction is

18 the addition of ozone across the double bond of OL forming an unstable primary ozonide (POZ).

19 Subsequently, the POZ decomposes, via two potential routes, through cleavage of the C-C bond

20 alongside one of the two O-O bonds. Both routes generate an aldehyde and an excited Criegee

21 Intermediate (CI) as products, but with differing chemical identity depending upon which O-O bond

22 is broken. In reaction route 1 (Scheme 1), nonanal (NN) and CI1 are formed; the CI1 can isomerise

23 to form stabilized azelaic acid (AA), a cyclic acyloxy hydroperoxide (CAHP), or octanoic acid

24 (OcA) and carbon dioxide (Gonzalez-Labrada, Schmidt et al. 2007). Alternatively, CI1 may be

25 scavenged via numerous potential association reactions with the co-produced aldehyde, other

26 carboxylic-functionalised moieties, (further) OL molecule double bonds, solvents or via self-

1   reaction. This extensive secondary chemistry reflects the reactivity of CIs (Zahardis and Petrucci,

2   2007). Similarly, in reaction route 2, oxo-nonanoic acid (ON) and excited Criegee Intermediate 2

3   (CI2) are formed, and CI2 decomposes by forming stabilized nonanoic acid (NA), or may be

4   scavenged via similar pathways as described above for CI1. The principal reaction products of OL

5   ozonolysis are therefore NN, AA, ON, and NA. There are also numerous reports in the literature of

6   the formation of higher molecular weight products (Hung et al., 2005; Reynolds, et al., 2006;

7   Zahardis et al., 2005) including esters (Hung et al., 2010) and peroxides (Reynolds et al., 2006;

8   Vesna et al., 2009; Zahardis et al., 2005; 2006; Ziemann, 2005). Although OL is essentially

9   hydrophobic, the reaction products are hydrophilic (Lee et al., 2012).
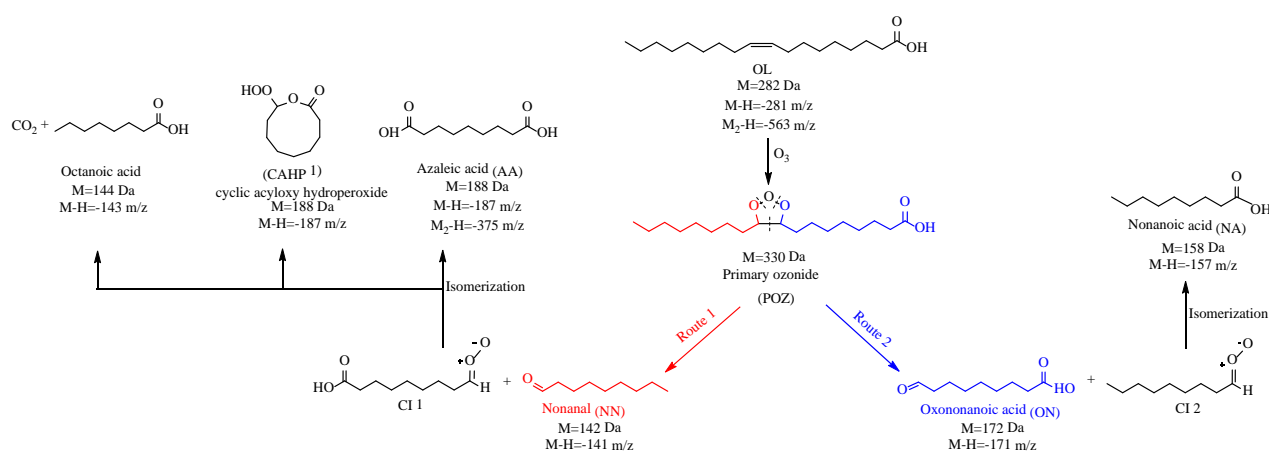
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11   **Scheme 1:** Initial steps and primary products of Oleic Acid (OL) oxidation by O_3.

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13   This study investigates the chemical mechanisms of OL ozonolysis and the resulting product

14   distribution, assessed as a function of particle size using a variety of novel online physical and

15   chemical analysis methodologies. The effect of relative humidity upon the (size dependent)

16   predominant reaction mechanism is also explored, and the resulting implications for the

17   atmospheric processing of organic aerosol are considered.

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1 **2.   EXPERIMENTAL**

2 **2.1   Experimental Design**

3 A schematic diagram of the experimental setup for the ozonolysis of OL aerosol is shown in Figure

4 1. The system comprised a gas handling and control system, an aerosol generator, ozone generator,

5 humidifier (bubbler), and an aerosol flow tube (AFT), coupled to aerosol particle characterisation

6 (SMPS - Scanning Mobility Particle Spectrometer - and ATOFMS - Aerosol Time-of-Flight Mass

7 Spectrometer) instruments, and monitors to measure ozone, temperature and relative humidity

8 (RH). The setup was designed so the ozonolysis of OL aerosol could be studied under well-defined

9 conditions within the AFT.

10

11 The AFT reactor consisted of a cylindrical tube formed from Pyrex (internal diameter = 10 cm,

12 length = 100 cm) which was sealed with aluminium flanges. The reactor was held in a vertical

13 orientation within a supporting frame.   Gas inlets and outlets were coupled to the flanges to allow

14 for the entrance and exit of the gas and aerosol streams. The ozone monitor, SMPS and ATOFMS

15 sampling outlets were located immediately downstream of the AFT.   Typical gas flows entering the

16 AFT were as follows : OL aerosol containing flow (1.0 slm in synthetic air), ozone containing dry

17 synthetic air (0.5 slm), and dry or humidified synthetic air (2.0 slm) resulting in a total flow rate of

18 3.5 slm which corresponded to a (plug flow) residence time of 135 s. In some experiments a

19 nitrogen flow was used to generate the aerosol (via homogeneous nucleation of OL); in these cases

20 a small flow of oxygen was added to ensure the air composition within the AFT had an atmospheric

21 $N_2:O_2$ ratio. Constant gas flows were achieved using calibrated mass flow controllers. The

22 Reynolds number for the AFT was 13.1 and hence flow within the reactor was predicted to lie

23 firmly within the laminar regime after an initial flow development length (estimated at 4.6 cm, i.e. <

24 5% of the total length).

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2 To ensure that particle processing was dominated by the AFT section of the setup, the additional

3 residence time within the sampling and characterisation stages was assessed. The calculated

4 residence time of the aerosol in the sampling tubes (<2 s) and SMPS (<1 s) (von Hessberg et al.,

5 2009) was negligible. Within the ATOFMS the residence time is very small, and the low operating

6 pressures within the ATOFMS (~10^{-7} Torr) minimize interactions between ozone and particles, and

7 hence particle reactivity is considered to be terminated at the very early stages of the ATOFMS

8 sampling nozzle.

9

10 A hygrometer probe positioned inside the AFT was used to monitor RH and temperature.

11 Experiments were typically run under both dry and humidified conditions which corresponded to

12 measured relative humidities of 0.5±0.02 % and 65.0±0.2%, respectively.

13

14 Each experiment was run back-to-back with and without the presence of ozone, with all other

15 conditions remaining identical, to allow measurement of both the initial OL particles before

16 ozonolysis and the resulting aerosol post ozonolysis. Measurements were conducted once the

17 aerosol concentration, size distribution and ozone concentration had stabilised. The system was

18 flushed between runs to prevent contamination from prior experiments.

19

20 **2.2   Scanning Mobility Particle Sizer (SMPS)**

21 To follow changes in the OL particle size distribution due to ozonolysis, an SMPS system

22 comprising an electrostatic classifier (TSI model 3080), neutralizer (TSI model 3076), long column

23 differential mobility analyzer (LDMA, TSI Model 3081) and condensation particle counter (CPC,

24 3022A) was used to monitor the aerosol size distribution. A particle density equal to that of pure OL

25 (0.891 g cm^{-3}) was applied to determine changes in the number, diameter, volume and mass

26 distribution of OL particles. Aerosols were monitored in the size range of 15.1 – 667 nm, every 3

1 minutes, before and after exposure to ozone. For each experimental setting, a minimum of 10 sizing

2 scans were averaged.

3

4 **2.3   Aerosol Time of Flight Mass Spectrometer (ATOFMS)**

5 A detailed description of the ATOFMS is given elsewhere (Gard et al., 1997; Sullivan and Prather,

6 2005; Dall'Osto et al., 2006); briefly, the ATOFMS is an on-line single particle instrument which

7 was applied here to measure the physical and chemical characteristics of OL aerosol before and

8 after ozonolysis. The ATOFMS deployed in this study (TSI, Model 3800-100) sampled particles in

9 the size range 100 – 3000 nm. For analysis of individual particle composition, time of flight mass

10 spectrometry (TOF-MS) was used with a laser (266 nm) desorption ionization (LDI) source.

11

12 The ATOMFS sampled from the AFT at a flow rate of 0.3 slm. When an aerosol flow enters the

13 ATOFMS it is directed through an expansion nozzle and skimmers during which particles are

14 accelerated to a velocity characteristic of their aerodynamic size; the smaller the particle the higher

15 the speed. Velocities of individual particles, hence aerodynamic size, are measured via scattered

16 light from two CW timing lasers (532 nm) positioned a known distance apart. The particle velocity

17 determines the timing of the subsequent LDI pulse to allow for interaction between particles and

18 LDI. Once ionized, the molecular fragments from the particle are directed to both positive and

19 negative polarity time-of-flight mass spectrometers.

20

21 The detection of particles within the ATOFMS is reliant upon the LDI pulse generating ions,

22 requiring the particle to have sufficient absorption properties at 266 nm. We found that the

23 interaction between the LDI and both OL and its ozonolysis products produced only negligible ions,

24 failing to meet the detection criteria of ATOFMS. To increase the absorption properties of the

25 particles, the dye Nile Blue Sulphate (NBS) was introduced into the OL particles. NBS effectively

26 absorbs light at the lasing wavelength of the LDI laser with absorption peaks occurring at 624, 325

8

and 276 nm) and therefore its introduction into the aerosol provides a methodology for efficient ionization. NBS was also separately introduced into calibration particles formed from each of the primary products of OL ozonolysis (See Results section).

Although the ATOFMS response is matrix-sensitive and may not directly reflect the quantitative composition of the measured aerosol (Gross et al., 2000; Fergenson et al., 2001; Bhave et al., 2002; Sullivan and Prather, 2005; Allen et al., 2006; Dall'Osto et al., 2006), relative peak intensities can be meaningful when particle constitution within a measured sample has not changed considerably during a given investigation (Gard et al., 1998; Finlayson-Pitts and Pitts, 2000). In this sense, the use of ATOFMS as a semi-quantifiable approach for the analysis of aged OL particles as a function of particle size is justifiable because individual laboratory based particles studied under experimentally controlled conditions are far less complex - and variable - than atmospheric aerosol (Allen, 2004). Moreover, ATOFMS information on particle number and the abundance of different species by particle size are possible with high time resolution (Dall'Osto and Harrison, 2012). Most importantly, the relative signal intensities arising from the major oxidation products of the OL-O_3 heterogeneous reaction were studied using individual authenticated standards, the analysis of which was used to calibrate signals arising from reacted oleic acid particles. The minimal residence time of the ATOFMS instrument (compared with, for example, drift-tube techniques) minimises the scope for within-detector ion-molecule chemistry to contribute to the observed signals; further confidence in this regard is achieved through the focus upon changes in signal upon addition of the chromophore, and as a function of chemical reaction conditions (in comparison with a focus upon direct spectra).

The ATOFMS particle sizing was calibrated with standardized polystyrene latex spheres (PSL, calibrated diameters = 0.2, 0.3, 0.4, 0.56, 0.7, 1.3, 2, and 2.5 μm). Dilute PSL suspensions were prepared in deionised water and nebulised using a constant output atomiser and diffusion dryer.

1 Analysis of the measured velocity of the nebulised PSL particles as a function of their known

2 particle sizes provided a calibration curve for the ATOFMS particle sizing.

3

4 The mass scale of the ATOFMS was calibrated through nebulisation of two multi-element

5 commercial standards (see Chemicals and Reagents, below) and further standards prepared for this

6 study reflecting the principal products of the OL ozonolysis system : OL, AA, NN, NA and 4-

7 oxononanoic acid (4-ON) together with NBS (added chromophore). The obtained time of flight of

8 the resulting ions in each case was related to the corresponding expected m/z values. This approach

9 allowed for the ATOFMS response to be directly calibrated up to 563 Da. m/z values reported here

10 beyond this limit have been calibrated using the molecular masses for those higher molecular

11 weight products for which identities have been proposed. To assess the relative sensitivity of the

12 ATOFMS towards OL, NN, AA, NA and 4-ON, a standard mixture solution containing OL, NN,

13 AA, NA, and 4-ON at equimolar concentrations, plus traces of the laser dye, NBS, was prepared in

14 methanol and aerosolized using liquid atomisation. The ATOFMS datasets were all processed

15 using MS Analyse and Microsoft Access software.

16

17 **2.4   Generation and Measurement of OL Aerosol and Ozone**

18 Ozone was generated by flowing 0.5 slm of synthetic air through a quartz photolysis tube

19 illuminated by a mercury UV pen-ray lamp The ozone concentration was monitored downstream of

20 the AFT using an ozone monitor (2B Technologies, Model 205). The ozone mixing ratio was 20

21 ppm. An elevated O_3 level allows use of short contact times in the AFT and corresponds to an

22 integrated exposure approximately equivalent to about one day in ambient air [if potential within-

23 particle diffusion limitations may be neglected].

24

25 Two approaches were used to generate OL aerosol. For the physical characterisation of OL aerosol

26 using the SMPS, liquid OL particles were generated by homogenous nucleation of pure OL vapour.

1   The low-volatile OL liquid was heated in a Pyrex vessel, in an insulated and temperature-controlled

2   oven (120 ± 0.5 °C), to create OL vapour. To aid evaporation 0.5 slm of N_2 gas was bubbled through

3   the liquid OL. Subsequently, the vapour underwent homogeneous nucleation into OL aerosol

4   particles accelerated by the introduction of a 0.5 slm flow of N_2 gas at ambient temperature. The

5   resulting aerosol stream was passed through a reheating tube (200 °C) to homogenise the aerosol

6   and narrow the resulting particle size distribution. All steps in the procedure were kept at

7   temperatures significantly below the boiling point of OL (360 °C) to prevent pyrolysis. Overall,

8   these operating conditions generated OL aerosols with a geometric standard deviation σ_g of 1.2, a

9   mass median diameter close to 400 nm and a total particle number concentration of $\sim10^6$ cm^{-3}.

10

11   For the chemical product study of OL aerosol using the ATOFMS, a different method for

12   generating OL aerosol was required as a consequence of the necessity of doping the aerosol with

13   NBS dye (see discussion above). NBS was added at a low concentration (5×10^{-5} M) to liquid OL

14   and the mixture made up into 0.15 M solutions of OL in methanol. Subsequently, a constant output

15   atomiser (TSI, Model 3076) was used to generate a polydisperse aerosol ensemble from the

16   solution. The atomiser requires a constant flow rate of 3.0 slm for operation. Of this aerosol-

17   containing flow, 1.0 slm was directed into the AFT after it passed through a series of two silica gel

18   diffusion dryers (TSI: Model 3062) for the removal of methanol prior to entry to the AFT, while the

19   remaining flow was vented. Under the experimental conditions of this study, methanol is expected

20   to fully evaporate and indeed no methanol signal was detected in any of the ATOFMS mass spectra.

21   Generation of the other aerosols containing NBS (i.e. calibration standards) was performed in the

22   same manner.

23

24   **2.5   Chemicals and Reagents**

25   Liquid OL (purity > 99.0%), solid crystalline azelaic acid (AA, 98%), liquid nonanal (NN, 95%),

26   liquid nonanoic acid (NA, 96%) and liquid 4-oxononanoic acid (4ON, 95%) were purchased from

1 Sigma-Aldrich and used as supplied. For ATOFMS mass calibration, multi-element standard

2 solutions were used: standard A (900 μg/ml: Ba, Pb, Li, K, Na & V in 5% HNO_3) and standard B

3 (900 μg/ml: Ag and Mo in 5% HNO_3) obtained from TSI Inc. The size calibration of the ATOFMS

4 used polystyrene latex spheres (PSL, diameter range: 0.1 to 2.5 μm, Duke Scientific). The cationic

5 laser dye, bis[5-amino-9-(diethylamino)benzo[a]phenoxazin-7-ium] sulphate, Nile Blue Sulphate

6 (NBS), used to modify the spectroscopic properties of OL particles and laboratory grade methanol

7 were obtained from Fisher Scientific. Synthetic air and nitrogen free oxygen (N_2) (99.9% stated

8 purity) were supplied by BOC and purified by passing through charcoal traps (Grace Discovery

9 Science, Efficiency<20 ppb) to remove trace hydrocarbon impurities. Traces of water were removed

10 from the gas lines using a trap (Sigma-Aldrich, Molecular Sieve 5A Moisture Trap).

11

12 **3.   RESULTS**

13 **3.1   Effect of Ozonolysis on Particle Size**

14 The size distributions of ozonolysed and non-ozonolysed OL particles, under both dry and

15 humidified conditions, are shown in panels (a) and (b) in Figure 2. To aid comparison, the wet and

16 dry results are combined, and normalized to the peak signal of the non-processed particle, in Figure

17 2 panel (c). It can be clearly observed that OL particles lose mass upon oxidation under both dry

18 and humidified conditions. The modal diameter decreases from 400±14 to 346±13 nm and from

19 372±13 to 322 ±12 nm under the dry and humidified conditions, respectively, corresponding to a

20 total mass loss (estimated from the SMPS) of 32.5±5.3% and 26.3±2.5% respectively. Although the

21 observed mass loss upon oxidation is consistent with previously reported qualitative and

22 quantitative observations (Morris et al., 2002; Katrib et al., 2004; Katrib et al., 2005b; Dennis-

23 Smither et al., 2012b; Lee et al., 2012), in view of the measured uncertainties the difference

24 between the two experiments, under different RH conditions, is statistically insignificant. The lack

25 of hygroscopic growth of the initial OL aerosol is consistent with the observations of Dennis-

26 Smither et al. (2012a) and Hung et al. (2005).

1

2 The decrease in the mobility diameter of OL particles upon oxidation is likely due to the formation

3 of volatile products that partition from the particle to the gas phase. This is consistent with previous

4 studies that have identified gas phase products. In particular the highly volatile nonanal (NN) and

5 the semi-volatile nonanoic acid (NA) products have been previously identified (Moise and Rudich,

6 2002; Smith et al., 2002; Thornberry and Abbatt, 2004; Vesna et al., 2009; Lee et al., 2012). The

7 formation of low volatility products, which remain in the particle phase, and typically possess

8 greater densities (ρ) than OL ($\rho = 0.891$ g cm^{-3}), such as azelaic acid (AA, $\rho = 1.225$ g cm^{-3}) and 9-

9 oxononanoic acid (ON, $\rho = 1.019$ g cm^{-3}) may also have an influence.

10

11 The fact that less mass loss was observed under humidified conditions than under dry conditions

12 (although the difference was statistically insignificant), suggests that the presence of water might

13 have an effect on the oxidation mechanism. Such a conclusion has been reported in the literature for

14 similar oxidation conditions (Gallimore et al., 2011). It has been suggested that the presence of

15 particle phase water leads to different CI reactivity (i.e. an alternate CI fate) which results in the

16 formation of less volatile products (such as organic acid formation) and hence less particle mass

17 loss (Gallimore et al., 2011). This is surprising because OL and its oxidation products, even though

18 they are smaller and contain more hydrophilic chemical moieties (carboxylic acid and hydroxyl

19 groups), have low solubility in water. Results from other experimental studies indicate that OL has

20 mild hygroscopicity (Andrews and Larson, 1993; Kumar et al., 2003; Vesna et al., 2008)

21 Furthermore thermodynamic modelling indicates that OL and ozonolysed OL particles show small

22 but non-negligible hygroscopicity at RH 65% (Lee et al., 2012).   It should be noted that

23 hygroscopicity is defined by mass or diameter growth upon uptake of water referenced to the dry

24 state. Since the molar mass and molar volume of water is much smaller than OL and its oxidation

25 products, the molar ratio of water to OL or its oxidation products is much greater than the reported

26 hygroscopicity would suggest. Dennis-Smither et al. (2012a,b) report a reduction in particle size

1 with oxidation of OL and that hygroscopicity increases with oxidative aging. While Lee et al.

2 (2012) observed no change to the reaction scheme of OL with increased humidity, Vesna et al.

3 (2009) report changes in product yields.

4

5 **3.2   Chemical Characterisation of OL and Ozonolysed OL Aerosol**

6 Particles leaving the AFT were sampled by the ATOFMS. Ionisation within the instrument leads to

7 ions which enter two time-of-flight mass spectrometers. Both positive and negative mass spectra

8 were recorded for all detected particles. The negative mass spectra show a greater abundance of

9 peaks which is expected based upon the predicted reaction pathways (see Scheme 1) because most

10 OL oxidation products contain carboxylic acid and/or aldehyde moieties which are principally

11 detected as deprotonated molecular ions [M-H]$^-$ formed via proton abstraction. To allow easier

12 interpretation and minimise fragmentation, the LDI laser fluence of the ATOFMS was kept very

13 low (0.4–0.8 mJ per pulse) as compared to other studies (1.3–1.6 mJ) (Silva and Prather, 2000;

14 Dall'Osto and Harrison, 2012). Nevertheless, it should be noted that some fragmentation of ions in

15 the ATOFMS system is expected and this can lead to difficulty in interpretation.

16

17 **3.3   Mass Spectrometric Analysis of NBS and OL**

18 To probe for any influence of NBS on the chemistry of OL ozonolysis two preliminary tests were

19 carried out. Firstly the mass spectra of NBS aerosol with and without O_3 present were recorded

20 (Figure S1a). Secondly the mass spectrum of aerosol composed of NBS and OL without O_3 was

21 recorded (Figure S1b). NBS displays three strong signals in the positive spectrum at m/z +274 and

22 +308 assigned to the cationic dye fragments and +318 assigned for the non-fragmented dye cation.

23 The two strong signals in the negative spectra at m/z -96 and -80 correspond, respectively, to the

24 sulphate (SO_4^{2-}) and sulphite (SO_3^{2-}) ions of the dye. No changes in the NBS spectra were detected

25 when NBS particles were exposed to ozone (Figure S1a). When OL was added to NBS solution in

26 the absence of ozone, the only additional peaks observed were the deprotonated mass signals from

1   OL and the OL dimer. It should be noted that NBS peaks were always observed in all mass spectra

2   shown henceforth. However, for the ease of the analysis, all NBS peaks are omitted from the

3   presented mass spectra.

4
5   **3.4   Mass Spectrometric Analysis of OL Oxidation Products**

6   In order to quantify the products of OL ozonolysis the relationship between the mass spectral signal

7   and the concentration needs to be quantified. Standards for OL, AA, NN and NA were all

8   commercially available. However, ON was not commercially available but its isomer 4-

9   oxononanoic acid (4-ON) was. The detection of ON via the ATOFMS is expected to be near-

10   identical to that of 4-ON. ON contains a carboxylic acid and an aldehyde group whereas 4-ON

11   contains carboxylic acid and ketone groups. Both acids, derived from ON and 4-ON, are expected

12   to exhibit a deprotonated carboxylate ion $[M-H]^-$ and were detected in the negative ion mass

13   spectrum at m/z -171. Standards for the high molecular weight oxidation products are not available

14   commercially and are difficult to synthesise and hence were unavailable for this study.

15

16   Figure 3 shows the positive and negative mass spectra of the individual standards. Unreacted OL

17   aerosol (282 Da) was seen as a deprotonated OL molecular ion at m/z = -281 and as a singly

18   deprotonated dimer at m/z -563. The deprotonated molecular ions of AA (MW = 188 Da), NA (158

19   Da), NN (142 Da) and 4-ON (172 Da) are seen as major peaks at m/z -187, -157, -141 and -171

20   respectively, while the singly deprotonated dimers of AA and NA were measured at m/z -375 and -

21   315, respectively. In the positive ion mass spectra of AA a signal at m/z +155 was attributed to the

22   fraction of the molecule after the loss of an HO_2 fragment $[M-HO_2]^+$. The molecular ion peak

23   measured at m/z +113 in the NA positive ion mass spectra was assigned to the molecular fragment

24   of NA after the loss of CO_2H, $[M-CO_2H]^+$. It is important to note that the ATOFMS intensity for

25   individual ions is dependent upon the total composition of the particle.

26

15

1 To investigate ATOFMS sensitivity towards OL and the primary ozonolysis products, a solution

2 containing an equimolar mixture of OL, AA, NA, NN and 4-ON was examined under the same

3 experimental conditions as the oxidation experiments. Figure 4 shows the results of the ATOFMS

4 (relative) sensitivity to the deprotonated molecular ions of the five component mixture. An average

5 spectrum comprised of 200 spectra was used for the analysis. It is evident that the ATOFMS

6 sensitivity varies considerably between the different compounds.

7

8 The figure indicates that the signals observed for the five component mixture is in the following

9 order: AA >> ON > OL > NA > NN. The much greater signal from AA compared to the other

10 species can be partially understood because it is a dicarboxylic acid and thus has a greater

11 possibility of being ionized to the carboxylate ion compared to the species containing only one

12 carboxylic acid moiety. ON and OL have similar signals which is expected since they both contain

13 one carboxylic acid group. The signal from NN is the smallest probably because it only contains an

14 aldehyde moiety which is more difficult to ionize compared to carboxylic acids. Previous studies on

15 NN detection using soft ionisation and ultra-high resolution mass spectrometry also showed weak

16 NN peaks in the mass spectra. The authors attribute this observation partly to the poor signal

17 produced from aldehydes by the deployed techniques (Grimm et al., 2006; Hosny et al., 2013). The

18 high volatility of NN will also lead to the partitioning from the aerosol to gas phase thereby

19 reducing the particle phase concentration, although it is noted that this process will also happen in

20 the processed oleic acid particles.

21

22 The relative peak areas of the five component mixture were converted into sensitivity correction

23 factors (SCF) that allow for relative concentrations of these components to be estimated within the

24 ozonolysed OL particles.   It should be noted that this approach – as in many other analytical

25 approaches – is not sensitive to any matrix effect of ozonolysed OL particles which might have an

26 impact of suppressing or enhancing ATOFMS signals of individual components. To minimize

1   possible matrix complications, the relative peak intensities of the individual components within

2   each measured sample were used to describe the component distribution of the chemically aged OL

3   polydisperse aerosol. As mentioned in Section 2.3, satisfactory ATOFMS measurement of the aged

4   OL particles required improving the efficiency of the LDI by optically modifying the matrix of OL

5   particles using added NBS laser dye. This modification resulted in a substantial enhancement in

6   ATOFMS detection of particles apparently by maximising the absorption efficiency of the pulsed

7   laser by the particles.

8

9   **3.5   Mass Spectrometric Analysis of Aged OL Particles**

10   The size binned ATOFMS data show distinct differences in composition between small particles

11   (D_p<0.3 µm) and large particles (D_p>0.3 µm). Small particles are characterised by the presence of

12   compounds with molecular weights lower than the parent OL molecules (m/z<282), whereas large

13   particles are characterised by significant formation of higher molecular weight (HMW) compounds

14   (m/z>282). Smith et al. (2002) measured ozone uptake by OL particles as a function of particle

15   size. They found that the effective uptake coefficient of ozone decreased with increasing particle

16   size, concluding that this resulted from the reaction being limited by the diffusion of OL within the

17   particle. Under such conditions, reactions between intermediate products and OL become

18   important, hence leading to the formation of higher molecular weight species (Hung et al., 2005;

19   2007; Katrib et al., 2004). Reactions between ozonolysis products and Criegee intermediates also

20   lead to high molecular weight products when the supply of ozone is limited (Reynolds et al., 2006;

21   Zahardis et al., 2006). Our interpretation of the mass spectral information gained from the

22   ATOFMS appears in Table 1.

23

24   **3.6   Mechanisms and Mass Spectrometric Analysis of Small Particles**

25   The mass spectrum of small oxidised OL particles (average of 300 spectra) is shown in Figure 5. In

26   addition to the parent OL signal measured at m/z -281, three of the primary ozonolysis products:

17

1 AA, ON, and NA, were observed at m/z -187, -171 and -157, respectively. Peaks measured at m/z=-

2 143, -329 and -439 could be indicative of octanoic acid (OcA) (144 Da), the primary ozonide

3 (POZ), (see Scheme 1) and the secondary reaction product of CI2 with OL (reaction R 1),

4 respectively. The unidentified mass spectrum peaks are likely due to fragments (daughter ion peaks)

5 or aggregates of ions. No useful insights were gained from the positive mass spectrum.

6

Addition to OL carboxylic moiety

M=440 Da
M-H=-439 m/z
M-HO$_2$=+423 m/z
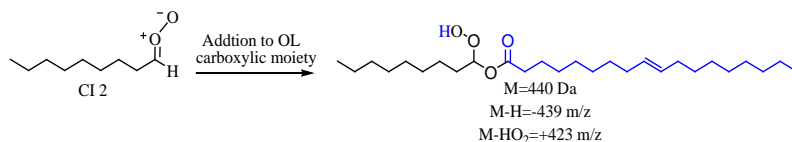
7   **R 1**

8

9 **3.7   Mechanisms and Mass Spectrometric Analysis of Large Size Particles**

10 The three frames, a, b, and c, in Figure 6 show negative and positive ion mass spectra of large ($D_p >$

11 0.3 μm) aged OL particles (average of 700 spectra). The negative ion mass spectrum is more

12 complex but compatible with the hypothesis of secondary chemistry and the formation of HMW

13 products. Generally, the spectrum demonstrates a strong peak of unreacted parent OL at m/z -281

14 and the appearance of the major oxidation products of ozone exposure. All four primary products

15 (AA, ON, NA and NN) are found to be in the particle phase including the highly volatile NN which

16 was not detected in the small particles. The presence of NN in large particles is consistent with the

17 observations of Dennis-Smithers et al. who studied supermicron sized particles (Dennis-Smither et

18 al., 2012b). The signals of AA and OcA which form through the molecular rearrangement of CI1,

19 were not detected in the larger particles, which might reveal that isomerization is predominant in

20 smaller particles and that it is the only formation source for these products. Both negative and

21 positive average mass spectra are characterized by the presence of HMW products presumably

22 formed by secondary association reactions of primary reaction products.

23

24 Previous studies have reported evidence of secondary chemistry occurring within the ozonolysis of

25 OL particles (Katrib et al., 2004; Hearn et al., 2005; Ziemann, 2005; Reynolds et al., 2006; Zahardis

18

1   et al., 2006; Zahardis and Petrucci, 2007; Lee et al., 2012; Hosny et al., 2013). The presence of the

2   liquid condensed phase substrate for the CIs minimizes their molecular rearrangement (Katrib et al.,

3   2004) via a solvent cage effect (Park et al., 2006) and maximizes their lifetimes. Therefore, reaction

4   probability of CIs with their corresponding carbonyl compounds to form secondary ozonide (SOZ)

5   or with the alkene functionality becomes more significant (Neeb et al., 1998; Moise and Rudich,

6   2002; Zahardis et al., 2005).

7

8   Table 1 proposes molecular structures for the observed HMW mass spectral signals based upon

9   polymerization mechanisms previously proposed in the literature (Smith et al., 2002; Hearn and

10   Smith, 2004; Katrib et al., 2004; Hung et al., 2005; Zahardis et al., 2005; Reynolds et al., 2006;

11   Zahardis et al., 2006; Hung and Ariya, 2007; Zahardis and Petrucci, 2007). For instance, the peak at

12   m/z -297 can be assigned for two isomeric compounds, 9-oxooctadecanoic acid (Ox1) or 10-

13   oxooctadecanoic acid (Ox2), depending on the type of the CI formed and the geometry of the CI

14   addition across the double bond of OL to form the C_{27} peroxide which can cleave to yield the

15   primary products ON or NN. Similarly, the detected signal at m/z -327 is determined as a reaction

16   product between the OL double bond and CI1 to form 9-oxooctadecanedioic acid (Ox3). Scheme 2

17   illustrates a proposed mechanism and product structures formed as a result of the reaction between

18   CI1 and the alkene functionality of OL. The experimental and modelling studies by Wang et al.

19   have shown that the addition of CIs across the double bond of OL is the dominant loss route of CIs

20   in OL-O_3 heterogeneous reaction system (Wang et al., 2016). We note that other reactions may also

21   contribute to the observed products - for example, fates of the OL dimer - which these experiments

22   cannot definitively constrain.

23

24   The presence of peaks correspond to Ox1 or Ox2 in the OL-O_3 heterogeneous reaction has been

25   reported previously (Hearn and Smith, 2004; Zahardis et al., 2006; Hung and Ariya, 2007) while the

26   reaction mechanism and product structures were first hypothesized by Katrib and co-workers

1   (Katrib et al., 2004).
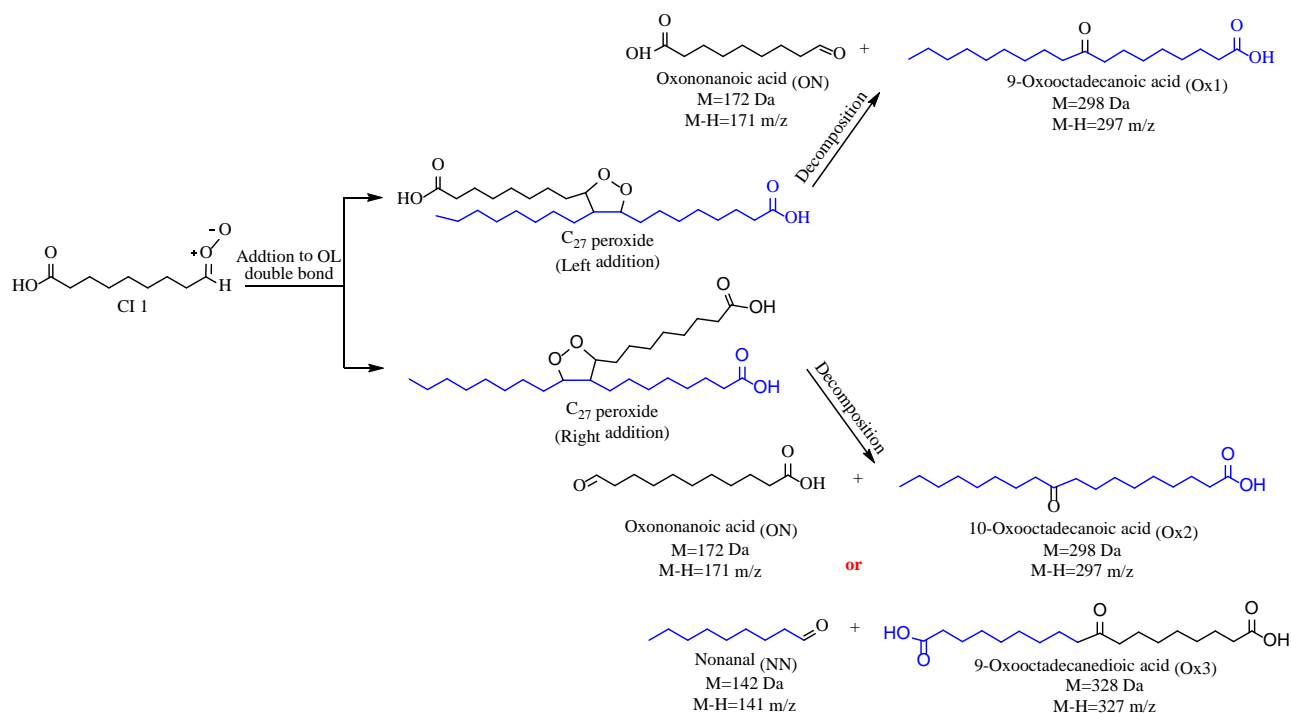
2

3   **Scheme 2:** Potential mechanism and product structures formed as a result of the reaction between
4   CI1 and the alkene functionality of OL.
5

6   The existence of the carboxylic acid functionality in the oxidation products offers reaction sites for

7   other CIs. The measured negative ion peaks at m/z -1019, -1189, and -1207 provide more evidence

8   of the incorporation of Ox1 or Ox2 as linear polymerization propagators. However, the ion peak

9   measured at m/z -644 is consistent with the formation of the ozonolysis product resulted from the

10   reaction of two ions of CI1 and the two carboxylic moiety of one molecule of Ox.

11

12   The addition of CI2 terminates the polymerization reaction as the $-CH_3$ group cannot react further

13   (Hung et al., 2005). The major HMW products ions at m/z -1025, -1079, -1214, -1292, -1310, -

14   1346, -1438, and -1524 may also correspond to polymerization products of Ox3 joined with other

15   propagators. Products related to the secondary reaction of ON are also observed in a number of

16   peaks. The negative ion at m/z -341 is most likely due to the combination of ON and CI1 to form

17   AAHP (azelaic acid hydroperoxide) with a loss of one molecule of water. Although the same

18   combination can lead to the formation of the SOZ, in such an arrangement a water molecule cannot

1   be lost. The in situ dehydration of AAHP was first observed and described by Zahardis and co-

2   workers (Zahardis et al., 2006). The resultant proposed structure suggests possible additional

3   moieties linking from molecule ends, the aldehyde and carboxylic acid groups, thereby growing

4   into a HMW linear polymer; Scheme 3 shows the overall reaction; although the intermediate

5   association complexes are unlikely to be detected in reality, with conversion to the M-H = 341

6   product and elimination of water proceeding as a concerted reaction.
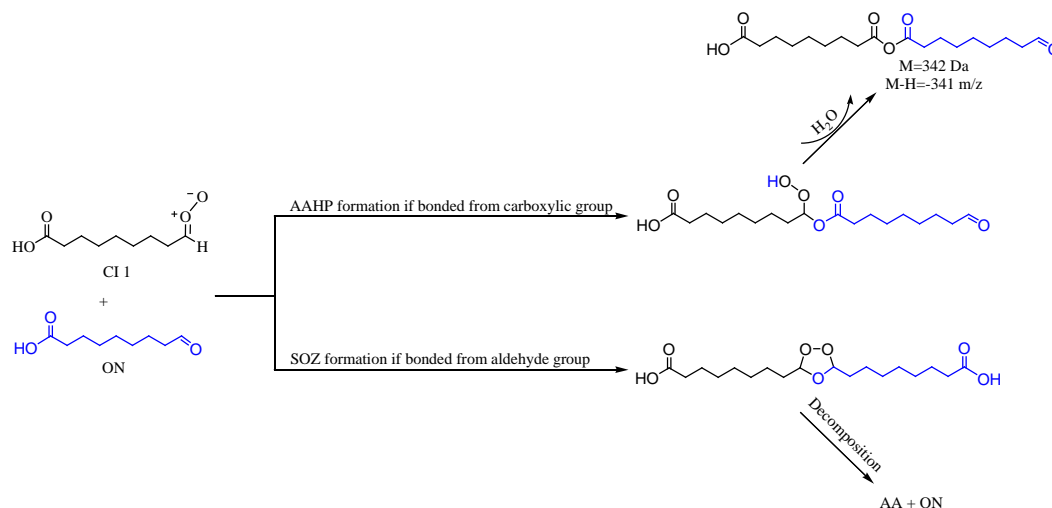
7
8   **Scheme 3.** Suggested reaction pathways and products of the secondary reaction between CI1 and
9   ON.
10

11   **3.8   Particle Composition as a Function of Particle Size**

12   The ATOFMS technique allows for particle composition data to be collected as a function of

13   particle size, thereby permitting the size dependence of the chemical aging of aerosols to be

14   determined. This section investigates how the composition of aged OL particles changes, under dry

15   and humidified conditions, as a function of particle size. The analysis only investigates the

16   following compounds which have clear MS peaks: OL, NN, AA, ON, NA and HMW compounds

17   (m/z > 282). The conversion from measured peak area to molar concentration is achieved using the

18   standard calibration factor (SCF). Since an accurate SCF for the HMW compounds is difficult to

19   determine due to the lack of standard laboratory calibrants, an upper limit value for the HMW SCF

20   was estimated using a mass balance approach. In particular, the ratio between the OL signal to the

21   total peak signals of the four primary oxidation products (AA, NN, NA and ON) was used to

1 estimate the ratio between unreacted OL and the products. The obtained ratio of each product was

2 corrected using the SCF to estimate the absolute molar ratio and hence the composition of the

3 primary products in the particles. The reduction in the molar ratio between the four products and

4 unreacted OL was used to estimate the SCF for the HMW (1 mole of OL produces 2 moles of the

5 primary products (Scheme 1)).

6

7 **3.9   Composition of Aged OL Particles Under Dry Conditions**

8 Figure 7 provides the compositional analysis, by applying SCFs to the raw ATOFMS data, of 1000

9 aged OL particles that were ozonolysed under dry conditions. The size dependent composition is

10 given in Figure 7a and the average composition of the aged OL polydisperse aerosol is shown in

11 Figure 7b. The average composition was achieved by taking the mean of all mass fractions

12 analysed regardless of size; since most particles were in the size range 0.3-0.5 µm this analysis will

13 be biased towards this size fraction. The overall size distribution of the analysed particles was in

14 the range 0.2-2.1 µm as measured by ATOFMS, see Figure 7c.

15

16 It can be clearly observed that the composition of aged particles is highly dependent on particle

17 size. Firstly, larger particles contain more unreacted OL than smaller particles. This can be

18 understood in terms of OL diffusivity within the particle and ozone flux into the particle. As the

19 reaction progresses and oxidation products evolve at the reacted surface layer, both OL diffusivity

20 within the particle and ozone diffusion to the bulk are reduced. Smith et al. (2002) reported a low

21 decay of OL in larger particles which was attributed by the limited OL diffusivity within the large

22 particles. However, OL molecules in smaller particles (at a diameter size of 200 nm and less) are

23 homogeneous in concentration and thus particles at this size and below are well mixed (Smith et al.,

24 2002; Pfrang et al., 2010) facilitating the oxidation process and therefore the smaller the portion of

25 unreacted OL. Additionally, in a simulation study, Shiraiwa et al. (2010) showed that ozone uptake

26 drops upon reaction, attributed to a decrease in OL concentration at the particle surface (while it

1   remains constant in the inner particle bulk, due to rapid bulk diffusion). It is therefore reasonable to

2   conclude that the larger the particle the less the molar ratio of O_3:OL in the overall particle phase

3   and hence the smaller the degree of oxidation and the greater proportion of unreacted OL. The

4   fraction of NN in the particle tends to increase with particle size. The smallest size fraction

5   contained only negligible amounts of NN. The vapour pressure of NN is 49.3 Pa at 25ºC (Daubert

6   and Danner, 1989) and hence it is a highly volatile species which tends to partition to the vapour

7   phase. In previous studies NN has been identified as one of the major volatile products (Moise and

8   Rudich, 2002; Thornberry and Abbatt, 2004; Zahardis and Petrucci, 2007; Vesna et al., 2009;

9   Dennis-Smither et al., 2012b), while few investigators detected NN in the particle phase using

10   supermicron (King et al., 2004) and supported millimetre sized particles (Hung and Ariya, 2007). It

11   is likely that NN has a low condensed phase yield in small particles due to its volatility that results

12   in fast partition to the vapour phase. NN signals, however, were frequently enhanced in large

13   particle spectra since its concentration in these particles was evolved from the secondary reaction

14   within the particles, and products in these particles might reduce the effective vapour pressure of

15   NN. Loss as vapour would also require diffusion within the condensed phase, which takes longer to

16   reach the surface in the larger particles. Wang et al. have argued that all observed NN in the particle

17   phase was formed via secondary reactions though the study was conduct on pure OL particles of

18   ~100 nm (Wang et al., 2016).

19

20   The mole fraction of AA, which arises from the rearrangement of CI1, was significantly larger in

21   small particles compared to larger particles. It diminished appreciably in the most reacted larger

22   particles suggesting the involvement of AA in the secondary chemistry associated more with larger

23   particles. NA which arises from CI2 showed different trend to AA, and was not detected in small

24   particles but observed in comparable fractions in larger size bins. This may reflect the volatile

25   nature of NA which would reduce its impact in secondary chemistry as it tends to partition away

1 from the particle. The steady decrease in ON mole fraction as particle size increases possibly

2 indicates the involvement of ON in the secondary reactions associated with larger particles.

3

4 Figure 7a suggests that reactions in larger particles can appreciably enhance the formation of HMW

5 (> 282 m/z) products. The chemical composition of larger particles, compared to smaller particles,

6 is characterised by more unreacted OL, increased HMW products and less AA and ON. The low

7 concentration of OL in small particles, however, minimizes these secondary pathways and

8 maximizes the formation of the SOZ which subsequently dissociates to form the four primary

9 oxidation products. The quantification analysis of processed OL particles by Katrib and co-workers

10 showed no trend for AA with increasing OL layer thickness (2-30 nm), but high ON yield as

11 compared to AA and NA yields was observed (Katrib et al., 2004), very consistent with our

12 findings. They attributed their finding to the reactivity of the AA and NA precursors, i.e. CIs, being

13 scavenged by OL. The presence of HMW products likely leads to greater aerosol viscosity and

14 hence lower diffusion rates within the particle (Hosny et al., 2016). Hence larger particles will

15 likely have slower diffusion rates which may in part explain the greater proportion of NN in the

16 larger particles. Moreover, measured viscosity of chemically aged aerosol might therefore be

17 expected to vary substantially with - increase with - particle size, of relevance to the interpretation

18 of aerosol particle viscosity measurements performed on individual super-micron samples.

19

20 **3.10   Composition of Aged OL Particles under Wet Conditions**

21 Figure 8 provides the corresponding compositional analysis, by applying SCFs to the raw ATOFMS

22 data, of 1000 aged OL particles that were ozonolysed under humidified conditions. Significant

23 differences in both the size dependent and average distribution of components between the dry and

24 humidified oxidation runs were observed.

25

The greatest differences between the dry and wet oxidation data, averaged over all size bins for the complete polydisperse aerosol ensemble, is the reduction in HMW products from 13.0% under dry conditions compared to 4.5% under humidified conditions. The proportion of unreacted OL increases under humidified conditions (29.7%) compared to dry conditions (25.3%). Whilst OL is negligibly hygroscopic, ozonolysed OL is slightly hygroscopic (Lee et al., 2012). This slight hygroscopicity suggests that particle phase water could act as a reactant. The differences in composition between the dry and the humidified experiments could be explained by the preferential reactivity of CIs with water molecules compared to OL thereby resulting in a lower OL consumption under wet conditions. This chemistry results in less destruction of additional OL molecules via the secondary pathways, and hence the higher portion of unreacted OL and the limited amount of HMW products which are observed concomitantly under wetter conditions. The size distribution data from the SMPS, which indicated less mass loss during the humidified oxidation, supports this argument. AA in small particle sizes showed a significant change in these experiments. The observed amount of AA in the smallest particles (300 nm) increased under humid condition (Figure 8 and 7). This result is partially supported by the study of Vesna et al. (2009), who used smaller particles (geometric mean diameter of 78 nm), and reported a similar trend with the smallest size bin of the particles in this study. The increase in the AA portion observed in response to the increase in the RH is consistent with our findings and it is attributed to water interference in supressing secondary reaction of AA with CIs, although Vesna et al. (2009) report a higher abundance of "unidentified products".

3.11 Kinetics of Oleic acid ozonolysis and product formation

The kinetics of the ozonolysis of a near-monodisperse sized ensemble of oleic acid particles in the size range of 0.4-0.5 µm with a mean diameter of 0.48 µm was investigated. Through the use of 3 different sized aerosol flow tubes (AFT), but with otherwise identical conditions to the experimental setup detailed previously, the interaction time between ozone and oleic acid particles could be

1   obtained at 20, 50 and 135 s. The longest time was the same as that used in the product distribution

2   study shown in Figures 7 and 8. In the absence of O_3 the effective interaction time is zero. We

3   observed that the concentration of oleic acid was negligible at the longest interaction time (135 s)

4   and hence the kinetics could only be followed in the time range of 0-50 s. The kinetics of the loss

5   of OL was measured under both dry (0.5% RH) and humidified (65% RH) conditions. The kinetics

6   of oleic acid loss appears to be a little faster under dry conditions compared to humid conditions.

7   This is surprising since oleic acid is only marginally hygroscopic and any water that is available

8   will likely partition to the hygroscopic region of the molecule around the carboxylic acid functional

9   group and not the lipophilic C=C double bond where the ozonolysis occurs.

10

11   The measured data, for both the dry and humid runs, show a linear relationship between the plot of

12    $\ln\{S(OL)/S(OL_0)\}$ versus interaction time, where S(OL) is the oleic acid signal, as shown in figure

13   9 below. This is consistent with surface limited reaction, as described by Case 3 kinetics initially

14   proposed by Hearn et al. (2005). Using the same approach as Hearn et al. (2005), using the Case 3

15   kinetics approach, to derive gamma uptake coefficients (γ), we obtain $\gamma_{dry} = 3.86 \pm 0.54 \times 10^{-4}$, γ_{wet}

16    $= 2.40 \pm 0.36 \times 10^{-4}$, and the value when both the dry and wet data are combined is $\gamma_{combined} = 3.13 \pm$

17   1.49×10^{-4}, where the stated errors are 2$\times$ the standard error of the mean. These values are similar

18   to previous measurements of γ, e.g. Hearn and Smith (2004), Moise and Rudich (2002), Hearn et al

19   (2005), Thornberry and Abbatt (2004), Ziemann (2005), Knopf et al. (2005). The observed kinetics

20   is what we expect if the initial reaction between OL and O_3 occurs at or near the surface. Steimer et

21   al. (2016) provide detailed mathematical analysis showing, for this reaction scenario, the

22   degradation rate of OL is size dependent but the uptake coefficient remains the same. It follows

23   that the kinetics of the ozonolysis of OL is surface limited rather than reaction diffusion limited,

24   detailed discussion of different kinetic regimes are provided in Berkemeier et al (2013).

25

1   In addition to the oleic acid reactive decay kinetics, we also observed the time dependent formation

2   of the four major first generation reaction products nonanoic acid (NA), azelaic acid (AA), nonanal

3   (NN) and oxononanoic acid (ON), see Figure 10. The high molecular weight (HMW) products

4   were not observable in these experimental runs which are consistent with the low levels observed in

5   figures 7 and 8 in the 0.3-0.5 µm size bin. It is noted that low levels of HMW products were

6   observed in the product distribution study but longer averaging times were used. It is clear that

7   whilst the kinetics of oleic acid loss is very similar under both dry and humidified conditions, there

8   are obvious differences in the formation kinetics of the four major first generation reaction

9   products. In particular, under dry conditions the reaction products form more promptly, and once

10   formed stay at relatively similar concentrations. Under humidified conditions the formation of the

11   peak concentration of the products is slower but also their subsequent loss is more substantial. The

12   kinetic data does not provide any definitive mechanistic understanding. However, these results are

13   consistent with the hypothesis that water can act as a reactant with the CI thereby reducing the

14   amount of secondary chemistry observed between OL and the primary reaction products, hence the

15   more stable product distribution after the initial ozonolysis step.

16

17   **3.11   Implications for Ageing Atmospheric Organic Aerosol**

18   The observed reactions could have consequences for the ability of OL-derived particles (and OA of

19   comparable functionality) to act as cloud condensation nuclei (CCN). Hygroscopicity greatly

20   enhances the ability of particles to act as CCN, and consequently, oxidised particles containing a

21   larger proportion of shorter-chain polar molecules are likely to be the most effective CCN for a

22   given particle size. King et al. (2009) use Köhler theory to demonstrate that the more oxidised

23   particles will activate at a lower supersaturation than unreacted OL particles. However, the

24   reduction in particle size which accompanies the loss of NN to the vapour phase will work in the

25   opposite sense of increasing the critical supersaturation due to an increased Kelvin effect, at least

26   for the smaller particles. However, since particles in the atmosphere are typically mixed and hence

1   unlikely to comprise purely OL, even when emitted, such a discussion is likely to be of very limited

2   relevance to atmospheric behaviour.

3

4   Heterogeneous reactions of organic particles directly alter the size, density and chemical

5   composition of the particles. These are the key parameters controlling the particle's lifetime in the

6   atmosphere and optical properties. While the size of the particle has considerable impact on the

7   deposition velocity, wet scavenging efficiency and scattering of light, the identity of the species

8   within the particle is the principal characteristic driving light absorption. The aging of organic

9   aerosol can result in the formation of light absorbing species. The large numbers of organic

10 functionalities in SOA such as carboxylate, hydroxyl, ketone and aldehyde groups may result in an

11 absorbing matrix that can exhibit optical properties dramatically different from those of parent

12 molecules. Particles containing light absorbers may lead to heating of the lower atmosphere

13 resulting in positive global radiative forcing. The impact of SOA on global radiative forcing is thus

14 one of the largest uncertainties in atmospheric science. On the other hand, the presence of absorbing

15 compounds in organic particles can stimulate photosensitization processes which might lead to

16 either reduction and oxidation of intermediates and products (Kolb et al., 2010). For the reasons

17 outlined above, because of the internally mixed nature of airborne particles, closer examination of

18 the properties of oxidised OL particles is unlikely to be of major relevance to prediction of the

19 properties of atmospheric aerosol.

20

21 **4.   CONCLUSION**

22 This study demonstrates a link between the particle size and reaction mechanisms within OL

23 aerosol. Aged finer particles are likely to be more hydrophilic due to the oxidation of the OL in the

24 particles and the formation of an early generation of more polar, and hence more hygroscopic

25 oxidation products. However, the combination of unreacted OL, HMW products and volatile

26 product observed in large particles suggests overall hydrophobicity of larger particles. The

1 difference in the reactivity of OL at different relative humidities with less OL destruction observed

2 under humid conditions can be explained by the preference of CIs for reaction with water molecules

3 over the reaction with OL, which results in less oligomerisation.

4

5 There has been some speculation in the literature as to the effects of oxidation upon atmospherically

6 relevant properties of OL particles, such as their ability to act as CCN. However, the relevance of

7 extrapolation to the atmospheric context, except in a generic sense, is extremely limited due to the

8 complex internal mixing of atmospheric particles, even when emitted, and hence the extreme

9 improbability of pure OL particles existing in the atmosphere. However, in our view, the value of

10 studies such as this is in the enhanced mechanistic understanding gained from treating OL oxidation

11 by ozone as a model system. In particular, there is limited understanding of processes leading to the

12 formation of highly functionalised oxidised high molecular weight products, such as those observed

13 in this work.

14

15 **ACKNOWLEDGEMENTS**

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TABLE LEGENDS

Table 1: A summary of proposed components and possible propagator combinations contributing to of observed mass spectral peaks corresponding to 44 oxidation products of the OL-O_3 system.

FIGURE LEGENDS

Figure 1: The experimental setup for the study of the heterogeneous oxidation of OL aerosol.

Figure 2: Particle mass size distributions for (a) pure and oxidised OL aerosol under dry condition (RH 0.5 ± 0.02%), (b) pure and oxidised OL aerosol under wet condition (RH 65.0 ± 0.2%) and (c) normalised particle size distribution of pure and oxidised OL aerosol under dry and wet conditions. Each curve represents the mean average of 10 measurements with accompanying standard deviation (σ).

Figure 3: Positive and negative ion mass spectra of OL, AA, NA, NN and 4-ON. Each spectrum presented represents the average from 100 mass spectra.

Figure 4: Relative peak area signals from aerosol particles generated from an eqimolar mixture of OL, AA, NA, NN and 4-ON. Peak areas are an average from 200 mass spectra.

Figure 5: Averaged ATOFMS negative and positive ion mass spectra for small processed OL particles (Dp < 0.3 µm).

Figure 6: Average mass spectra of dry aged OL particles (Dp>0.3 µm): (a) negative ion MS (b) zoom in plot of figure (a) and (c) positive ion MS.

Figure 7: ATOFMS data analysis of: (a) mole fraction of ozonolysed OL particles, under dry conditions, as a function of particle size, (b) total particle composition as given by mole fraction for all size fractions, and (c) Corresponding size distribution of the aged OL aerosol measured by ATOFMS.

Figure 8: ATOFMS data analysis of: (a) mole fraction of ozonolysed OL particles, under humidified conditions, as a function of particle size, (b) total particle composition as given by mole fraction for all size fractions, and (c) corresponding size distribution of the aged OL aerosol measured by ATOFMS.

Table 1: A summary of proposed components and possible propagator combinations contributing to of observed mass spectral peaks corresponding to 44 oxidation products of the OL-O_3 system

MW	- m/z	+ m/z	NN	OcA	AA or CI1	ON	NA or CI2	Ox1 or Ox2	Ox3	OL	Dehydration (-nH$_2$O)
142	141		1								
144	143			1							
158	157						1				
170	169				1						1
172	171					1					
188	187				1						
298	297							1			
314	313			1	1						
328	327							1			
342	341				1	1					
422	421	377*					1		1		1
440	439	423†					1		1		0
528	527	466*† 483* 547‡			3						3
644	643	599*					2	1			0
656	655				2		2				2
768	767				1		2		1		1
786	785	753♯			1		2		1		0
810	809	793† 777♯			2	1	2				3
844	843	811♯			3		2				2
864	863	819* 831♯			2	1	2				0
880	879				3		2				0
894	893	861♯			3	1	1				0
950	949	917♯			3		1		1		3
968	967	935♯			3		1		1		2
974	973				2		2		1		0
986	985	953♯			3		1		1		1
1002	1001				3		3				2

1020	1019	1003[†]			3		1	1			0
1026	1025				4				1		3
1038	1037	1005[⁒]			3		3				0
1051	1050	1070[‡]			4		2				1
1068	1067	1023[*] 1106[‡]			4		2				0
1080	1079				4				1		0
1162	1161	1112[†⁒]			3		2			1	0
1178	1177				3		4				1
1190	1189				4		1	1			1
1196	1195				3		4				0
1208	1207				4		1	1			0
1214	1213				5				1		3
1292	1291				3	1			1	1	3
1310	1309				3	1			1	1	2
1346	1345				3	1			1	1	0
1438	1437				6				1		1
1524		1458[⁒]			3		4		1		0

1   [*]M-CO$_2$H, [⁒]M-HO$_2$, [†]M-OH and [‡]M+H$_3$O.

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2   **Figure 1:** The experimental setup for the study of the heterogeneous oxidation of OL aerosol
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Figure 2: Particle mass size distributions for (a) pure and oxidised OL aerosol under dry condition (RH $0.5 \pm 0.02\%$), (b) pure and oxidised OL aerosol under wet condition (RH $65.0 \pm 0.2\%$) and (c) normalised particle size distribution of pure and oxidised OL aerosol under dry and wet conditions. Each curve represents the mean average of 10 measurements with accompanying standard deviation (σ)

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Figure 3: Positive and negative ion mass spectra of OL, AA, NA, NN and 4-ON. Each spectrum presented represents the average from 100 mass spectra

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Figure 4: Relative peak area signals from aerosol particles generated from an eqimolar mixture of
OL, AA, NA, NN and 4-ON. Peak areas are an average from 200 mass spectra

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4 **Figure 5:** Averaged ATOFMS negative and positive ion mass spectra for small processed OL
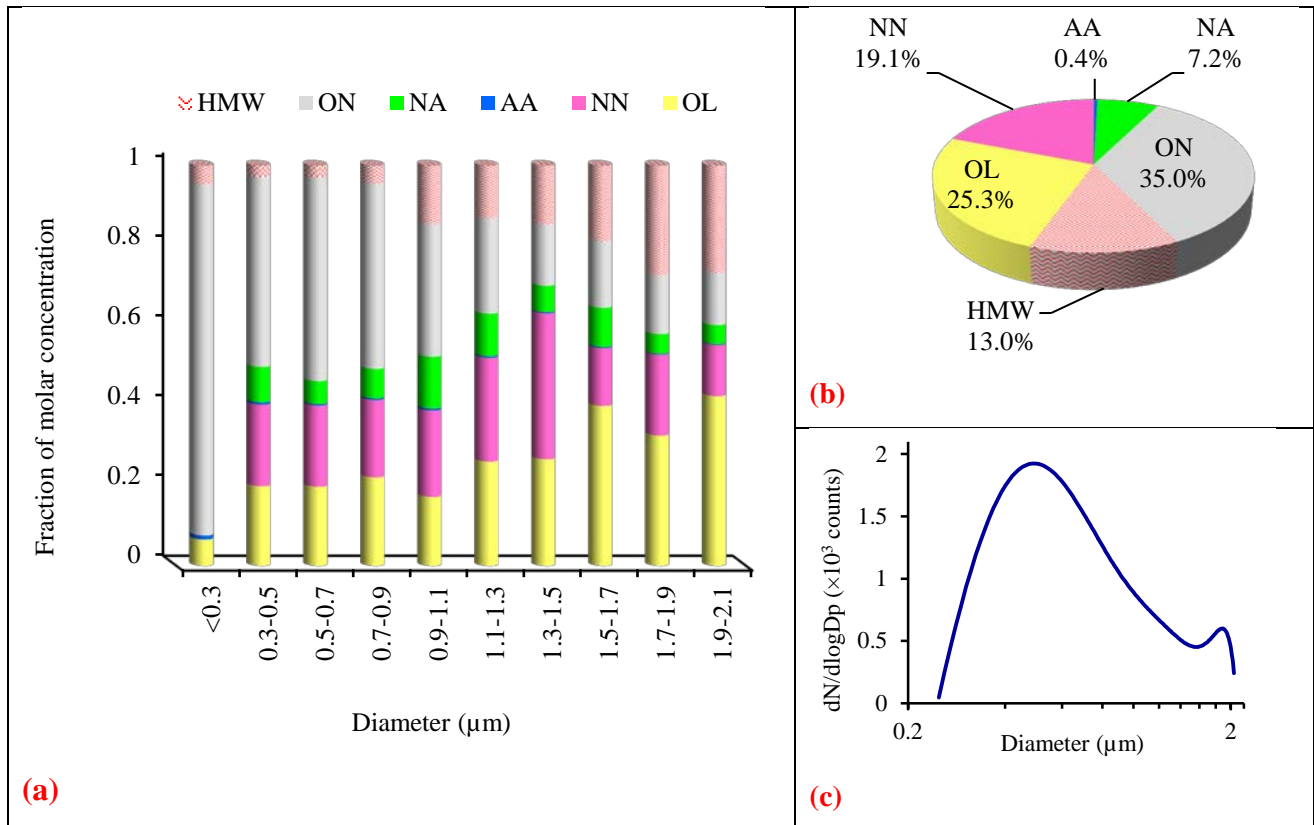5 particles (Dp < 0.3 µm)
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Figure 6: Average mass spectra of dry aged OL particles (Dp>0.3 µm): (a) negative ion MS (b) zoom in plot of figure (a) and (c) positive ion MS

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Figure 7: ATOFMS data analysis of: (a) mole fraction of ozonolysed OL particles, under dry
conditions, as a function of particle size, (b) total particle composition as given by mole fraction for
all size fractions, and (c) corresponding size distribution of the aged OL aerosol measured by
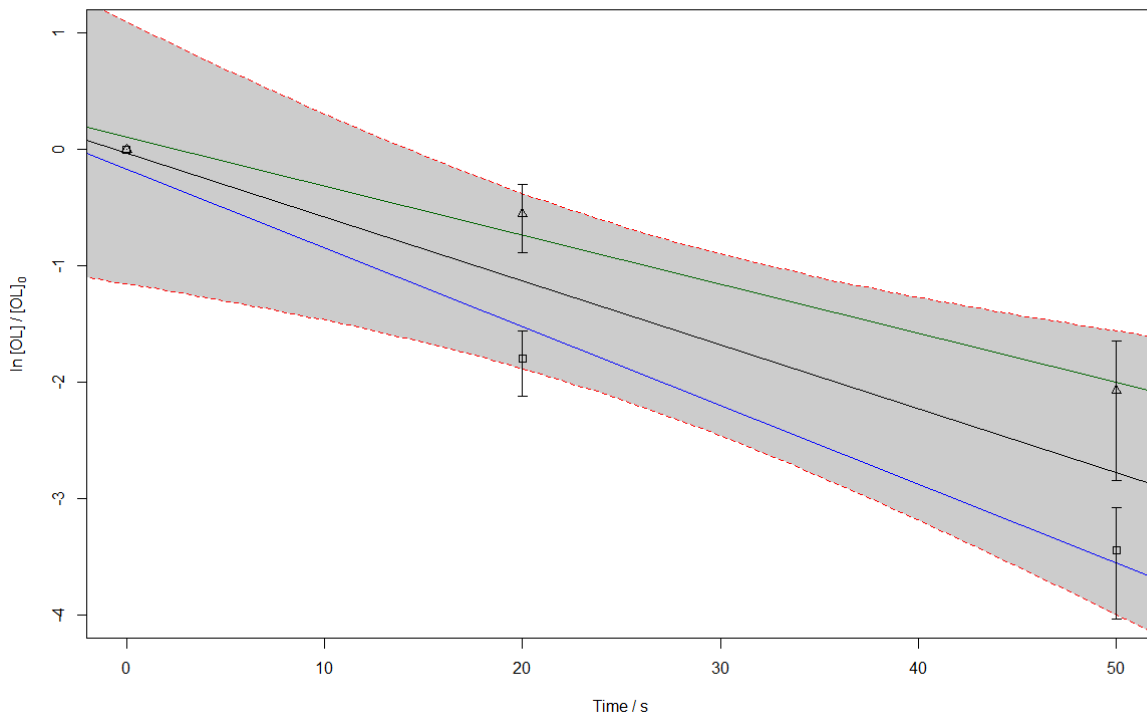ATOFMS

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Figure 8: ATOFMS data analysis of: (a) mole fraction of ozonolysed OL particles, under
humidified conditions, as a function of particle size, (b) total particle composition as given by mole
fraction for all size fractions, and (c) corresponding size distribution of the aged OL aerosol
measured by ATOFMS

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5 **Figure 9**. Reactive decay of oleic acid as a function of time for particles in the size range 0.4-0.5 µm (mean
6 diameter = 0.48 µm). Square symbols = measurements under dry conditions (RH = 0.5%), triangle symbols
7 = measurements under humid conditions (RH = 65%), black line = linear fit of combined dry and humid
8 data, blue line = linear fit of dry data set, green line = linear fit of humid data set, grey shading is the
9 confidence limits for the linear fit to the combined data set at the 95% confidence limits for prediction. Error
10 bars on the individual data points are 2× the standard error of the mean.
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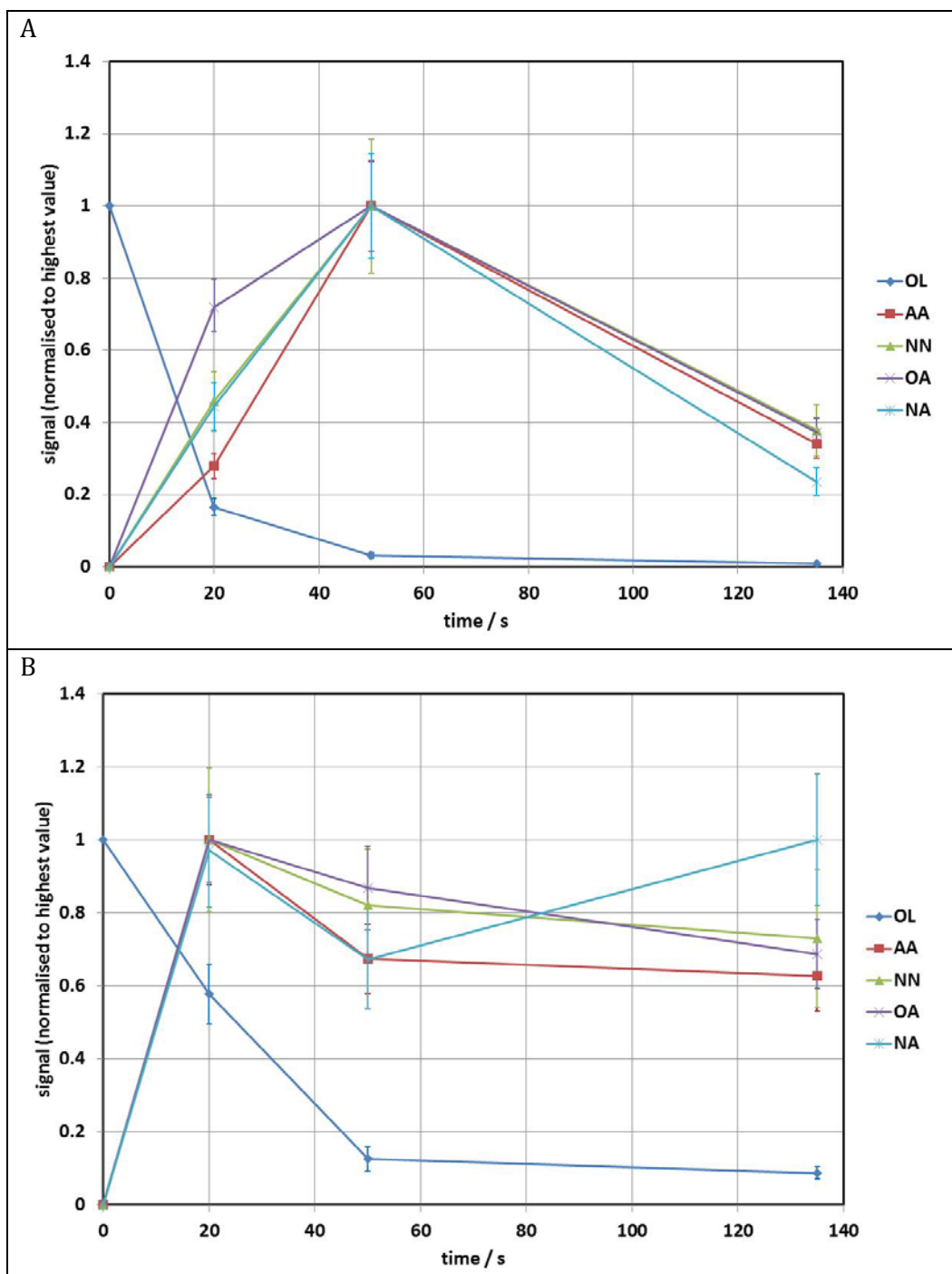
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Figure 10. Time dependent signals of oleic acid and the four major first generation ozonolysis
products. Graph A was obtained under dry conditions (RH = 0.5%) and B was obtained under wet
conditions (RH = 65%). To aid ease of comparison of the different time series, the signals for all
investigated species have been normalized relative to the peak signal achieved by the species
investigated. Error bars are 1× the standard error of the mean.
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