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# Secondary organic aerosol formation from photooxidation of naphthalene and alkylnaphthalenes: implications for oxidation of intermediate volatility organic compounds (IVOCs)

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1873

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## Abstract

Current atmospheric models do not include secondary organic aerosol (SOA) production from gas-phase reactions of polycyclic aromatic hydrocarbons (PAHs). Recent studies have shown that primary semivolatile emissions, previously assumed to be inert, undergo oxidation in the gas phase, leading to SOA formation. This opens the possibility that low-volatility gas-phase precursors are a potentially large source of SOA. In this work, SOA formation from gas-phase photooxidation of naphthalene, 1-methylnaphthalene (1-MN), 2-methylnaphthalene (2-MN), and 1,2-dimethylnaphthalene (1,2-DMN) is studied in the Caltech dual 28-m<sup>3</sup> chambers. Under high-NO<sub>x</sub> conditions and aerosol mass loadings between 10 and 40 μg m<sup>-3</sup>, the SOA yields (mass of SOA per mass of hydrocarbon reacted) ranged from 0.19 to 0.30 for naphthalene, 0.19 to 0.39 for 1-MN, 0.26 to 0.45 for 2-MN, and constant at 0.31 for 1,2-DMN. Under low-NO<sub>x</sub> conditions, the SOA yields were measured to be 0.73, 0.68, and 0.58, for naphthalene, 1-MN, and 2-MN, respectively. The SOA was observed to be semivolatile under high-NO<sub>x</sub> conditions and essentially nonvolatile under low-NO<sub>x</sub> conditions, owing to the higher fraction of ring-retaining products formed under low-NO<sub>x</sub> conditions. When applying these measured yields to estimate SOA formation from primary emissions of diesel engines and wood burning, PAHs are estimated to yield 3–5 times more SOA than light aromatic compounds. PAHs can also account for up to 54% of the total SOA from oxidation of diesel emissions, representing a potentially large source of urban SOA.

## 1 Introduction

Organic aerosol (OA) accounts for a large fraction of urban particulate matter (Seinfeld and Pankow, 2003; Zhang et al., 2007), and thus has important implications for health, climate and visibility. OA has traditionally been classified into two categories: primary organic aerosol (POA), which is directly emitted as particulate matter, and sec-

1875

ondary organic aerosol (SOA), which is formed from atmospheric oxidation of volatile organic compounds (VOCs). POA has been assumed to be nonvolatile and chemically inert in atmospheric models, emitted primarily from anthropogenic sources, such as diesel and gasoline engines, wood burning and cooking operations (Schauer et al., 1996). On the other hand, SOA, formed from gas-phase oxidation of biogenic and anthropogenic volatile precursors, is an important source of OA both on global and local scales (Goldstein and Galbally, 2007; Docherty et al., 2008). Nonetheless, the amount of OA observed in urban sites is consistently higher than that for which volatile SOA precursors can account (de Gouw et al., 2005; Heald et al., 2005; Volkamer et al., 2006). This POA/SOA classification has recently been called into question; during laboratory chamber photooxidation of diesel and biomass burning emissions, volatile SOA precursors, which consist primarily of light aromatic compounds such as benzene, toluene and xylene, can account for only a small fraction of the SOA produced (Robinson et al., 2007; Grieshop et al., 2008a,b). Robinson et al. (2007) proposed that gas-phase reactions of semivolatile compounds likely contribute to a large fraction of the SOA formed, as oxidation of these low-volatility organic vapors produces aerosol that closely resembles OA observed in field measurements (Weitkamp et al., 2007; Sage et al., 2008; Grieshop et al., 2008a). These semivolatile compounds, including long chain *n*-alkanes, polycyclic aromatic hydrocarbons (PAHs), and large olefins, have been termed intermediate volatility organic compounds (IVOCs).

PAHs account for a large fraction of semivolatile emissions from diesel engines and wood burning sources (Schauer et al., 1999a, 2001). While PAHs with fewer than 4 aromatic rings exist primarily in the gaseous phase, photooxidation of these compounds has been shown to yield high molecular weight (MW) oxygenated compounds (Sasaki et al., 1997; Bunce et al., 1997; Wang et al., 2007), which can partition into the particle phase and lead to significant SOA formation (Mihele et al., 2002). Moreover, nitronaphthalene, an oxidation product of naphthalene, and other nitro PAHs have been observed in urban particulate matter (Arey et al., 1989). The oxidation of PAHs in the gas phase will likely have significant impacts on both the quantity and the properties

1876

of SOA formed in urban atmospheres. In this study, we present results on SOA formation from gas-phase photooxidation of naphthalene, 1-methylnaphthalene (1-MN), 2-methylnaphthalene (2-MN), 1,2-dimethylnaphthalene (1,2-DMN). The SOA yields, the mass of SOA formed per mass of PAH reacted, are then used to evaluate the importance of PAHs as effective SOA precursors.

## 2 Experimental

Experiments were carried out in the Caltech dual 28-m<sup>3</sup> Teflon chambers. Details of the facilities have been described previously (Cocker et al., 2001; Keywood et al., 2004). Before each experiment, the chambers were flushed with dried purified air for >24 h, until the particle number concentration is <100 cm<sup>-3</sup> and the volume concentration is <0.1 μm<sup>3</sup> cm<sup>-3</sup>.

In most experiments, an inorganic seed aerosol was used to provide a surface for condensation of semivolatile oxidation products. Seed aerosol was generated by atomization of a 0.015 M aqueous ammonium sulfate solution. The aerosol size distribution, number and volume concentrations were measured with a differential mobility analyzer (DMA, TSI, 3081) coupled with a condensation nuclei counter (TSI, CNC-3760). The volume concentration was corrected for particle wall loss by applying size-dependent first-order loss coefficients, obtained in a separate seed-only experiment (Keywood et al., 2004).

After atomization of inorganic seed, the parent hydrocarbon was then injected into the chamber; the injection method varied depending on its phase under room temperature conditions. For naphthalene and 2-MN, the hydrocarbon was introduced into the chamber by flowing purified air through an FEP tube packed with the solid PAH at 1 L min<sup>-1</sup>. For 1-MN and 1,2-DMN, a known volume of the liquid PAH was injected into a heated glass bulb, and the vapor was carried into the chamber with 5 L min<sup>-1</sup> of purified air. The parent hydrocarbons studied and their stated purities are as follows: naphthalene (Sigma-Aldrich, 99%), 1-methylnaphthalene (Fluka, ≥97%),

1877

2-methylnaphthalene (Fluka, ≥95%) and 1,2-dimethylnaphthalene (Aldrich, 95%). The hydrocarbon was measured using a gas chromatograph with flame ionization detector (GC/FID, Agilent 6890N), equipped with an HP5 15 m×0.53 mm ID ×1 μm thickness column. The GC response was calibrated by dissolving a known mass of the parent PAH in dichloromethane, and vaporizing a small volume of that solution into a 38 L Teflon chamber.

For high-NO<sub>x</sub> experiments (initial [NO]>350 ppb), nitrous acid (HONO) was used as an OH precursor. Nitrous acid was prepared by adding 10 mL of 1 wt.% aqueous NaNO<sub>2</sub> dropwise into 20 mL of 10 wt.% sulfuric acid in a glass bulb. A stream of purified air was then passed through the bulb, sending HONO into the chamber. This process also forms NO and NO<sub>2</sub> as side products. The injection of HONO was stopped when [NO<sub>2</sub>] reached about 80 ppb in the chamber. Additional NO was added until total [NO] was about 400 ppb. NO<sub>2</sub> was monitored by a gas chromatograph with luminol detector (University of California, Riverside, CA), in which NO<sub>2</sub> and PAN were separated by gas chromatograph and detected by chemiluminescence of reaction with luminol (Burkhardt et al., 1988). Reaction of HONO with luminol is unlikely, and thus would not interfere with the NO<sub>2</sub> signal. A commercial chemiluminescence NO/NO<sub>x</sub> analyzer (Horiba, APNA 360) was used to monitor NO and total NO<sub>y</sub>. For low-NO<sub>x</sub> experiments, hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) was used as an OH precursor. Prior to atomization of the ammonium sulfate seed, H<sub>2</sub>O<sub>2</sub> was introduced by bubbling purified air through a 50% aqueous H<sub>2</sub>O<sub>2</sub> solution for 2.5 h at 5 L min<sup>-1</sup>.

After allowing for all concentrations to stabilize, irradiation was initiated. The temperature of the chamber reached 299 K, and the RH of the experiments was between 5 and 8%. Figure 1 lists the structures of the parent hydrocarbons and their OH reaction rate constants (*k*<sub>OH</sub>). The experimental conditions prior to beginning of irradiation and the results are summarized in Table 1. Photooxidation of 1,2-DMN under low-NO<sub>x</sub> conditions was not studied. Owing to its fast OH reaction rate, 1,2-DMN will likely be consumed near its ambient sources, where NO<sub>x</sub> concentrations are likely to be high.

A proton transfer reaction mass spectrometer (PTR-MS), custom modified Varian

1878

1200 system, was used to monitor the concentrations of various gas-phase species. Details of the instrument are given in Ng et al. (2007a). In brief, positively charged water clusters,  $(\text{H}_2\text{O})_n\text{H}^+$ , react via proton transfer with the analyte,  $A$ , to form the positively charged ion,  $A(\text{H}_2\text{O})_x\text{H}^+$ , where  $0 \leq x \leq n$ . Species were predominantly observed at  $m/z = M+1$ , where  $M$  is the molecular mass of the species. Hydrates,  $A(\text{H}_2\text{O})\text{H}^+$ , were also observed for formylcinnamaldehyde and phthalic anhydride.

Real-time aerosol mass spectra were obtained using an Aerodyne high-resolution time-of-flight aerosol mass spectrometer (HR-TOF-AMS) (DeCarlo et al., 2006). These data were also used to calculate the density of SOA by comparing the particle mass distribution obtained using the particle time-of-flight mode and the volume distribution obtained by the DMA in nucleation (seed-free) experiments (Bahreini et al., 2005).

### 3 Results

#### 3.1 Concentrations of NO and NO<sub>2</sub>

Figure 2 shows mixing ratios of NO and NO<sub>2</sub> in a typical high-NO<sub>x</sub> experiment (6/16). Following the onset of irradiation, photolysis of HONO produces OH and NO. The level of NO remained above 400 ppb, and no ozone formation was observed. Also, owing to the high NO levels, the reaction of RO<sub>2</sub> radicals (produced in the OH oxidation) with NO likely dominated over that with HO<sub>2</sub> or RO<sub>2</sub> radicals.

Following the addition of OH to the aromatic ring, the aromatic-cyclohexadienyl radical can react with NO<sub>2</sub> to form nitronaphthalenes or with O<sub>2</sub> to form peroxy radicals (Atkinson et al., 1987; Nishino et al., 2008). The relative rates of these reactions depend on the concentration of NO<sub>2</sub>, which remained between 50 and 100 ppb in all high-NO<sub>x</sub> experiments, as measured by the GC/NO<sub>2</sub>-PAN analyzer. This level is typical of urban polluted conditions, and suggests that the branching ratios in these experiments are relevant for regions where PAHs are commonly emitted. Nishino et al. (2008) reports that the reaction rates of OH-naphthalene adduct with NO<sub>2</sub> and O<sub>2</sub> are equal at

1879

[NO<sub>2</sub>] ≈ 60 ppb.

In the low-NO<sub>x</sub> experiments, NO and NO<sub>2</sub> concentrations were below limits of detection (2 ppb). Ozone was formed at a rate of 2.5 ppb hr<sup>-1</sup>, likely due to reactions of NO<sub>y</sub> contaminants on the chamber walls. Based on the low reaction rate constants of aromatic compounds with O<sub>3</sub>, such reactions are not expected to be significant.

#### 3.2 Gas-phase composition

The gas-phase photooxidation mechanism of naphthalene has been studied extensively (Sasaki et al., 1997; Bunce et al., 1997; Wang et al., 2007). Oxidation products identified in the current work by PTR-MS are consistent with those previously observed. Gas-phase mass spectra measured by PTR-MS during photooxidation of naphthalene under high- and low-NO<sub>x</sub> and the proposed structures for the ions observed are shown in Figs. 3 and 4, respectively.

#### 3.3 Aerosol formation

Figure 2 shows the mixing ratio of naphthalene and the mass concentration of SOA formed in the experiment on 6/16. Under high-NO<sub>x</sub> conditions, owing to the efficient photolysis of HONO, the average OH concentrations are on the order of  $1.5 \times 10^7$  molec cm<sup>-3</sup>, as estimated from the parent hydrocarbon decay. After 100 min of irradiation, HONO was depleted and, for naphthalene, further SOA formation was insignificant. Due to the relatively slow reaction rate of naphthalene with OH, only 85% of the initial naphthalene had been consumed before HONO was depleted. In one high-NO<sub>x</sub> experiment with an initial naphthalene concentration of 17 ppb (6/20), 4 ppb remained after HONO was depleted. More HONO was added to consume the remaining naphthalene, resulting in further aerosol formation. For the other PAHs, less than 7% of the initial concentration remained unreacted. Under low-NO<sub>x</sub>, a continuous supply of OH radicals was generated from the slow photolysis of H<sub>2</sub>O<sub>2</sub>, and the parent hydrocarbon concentration was below detection limits at the end of the exper-

iment. The average OH concentrations for these experiments were on the order of  $3 \times 10^6 \text{ molec cm}^{-3}$ .

The SOA mass concentrations and yields for all experiments are summarized in Table 1. To calculate the mass concentration of SOA, the volume concentration measured by the DMA was multiplied by its respective SOA density, obtained in separate nucleation experiments. Values for the densities are summarized in Table 2.

## 4 Aerosol formation

### 4.1 Relative rates of oxidation steps

Figure 5 shows the time-dependent growth curves, mass concentration of SOA formed,  $\Delta M_0$ , as a function of mass of hydrocarbon reacted,  $\Delta HC$ , for all experiments. For naphthalene under both high- and low- $\text{NO}_x$  conditions, the growth curves overlap; the lack of a vertical portion indicates that SOA formation likely results either from first-generation products, or from higher generation products that are formed rapidly (Ng et al., 2006; Chan et al., 2007). In the high- $\text{NO}_x$  experiment in which a second HONO addition was conducted (6/20), SOA growth from the second injection of HONO was identical to that in another experiment with a higher initial hydrocarbon concentration. These observations suggest that the first oxidation step is rate-limiting, and the trend in aerosol growth for naphthalene can be represented by a single growth curve.

On the contrary, the growth curves for methylsubstituted naphthalenes display a vertical portion (hook) at higher  $\Delta HC$ . In photooxidation of 1-MN and 2-MN under both high- and low- $\text{NO}_x$ , and that of 1,2-DMN under high- $\text{NO}_x$ , delayed aerosol growth was observed after a large fraction (>90%) of the parent PAH had been consumed. In addition, the degree of divergence for different growth curves is greater for 1,2-DMN, which has a larger number of methyl groups. While one cannot rule out the possibility of forming low-volatility compounds from photolysis of aromatic carbonyls and hydroperoxides, the delay in aerosol growth relative to the first oxidation step can also

1881

be explained by the relative rates of reaction. Figure 6 shows the oxidation mechanisms for naphthalene and 1-methylnaphthalene that lead to the major dicarbonyl product. As the aromatic ring becomes more substituted, the reaction rate of OH with the parent PAH increases (Phousongphouang and Arey, 2002). Further reaction of the dicarbonyl can occur by OH addition to the olefinic double bond, or by abstraction of aldehydic hydrogen atom. OH addition to the olefinic double bond is likely unaffected by substitution of methyl groups, and H-abstraction is slower for the methyl-substituted dicarbonyl product because of one fewer aldehyde group. As a result, the relative rate of the second step is slower for methyl-substituted naphthalenes, resulting in divergence of growth curves and delayed SOA growth relative to the oxidation of the parent PAH. Since further oxidation of gas-phase intermediates will likely lead to even lower volatility products and more SOA formation, the observed yields measured in these experiments represent a lower limit for PAH oxidation in the atmosphere.

### 4.2 $\text{NO}_x$ dependence of SOA growth

In this study, the SOA yields under high- and low- $\text{NO}_x$  conditions are 25–45% and 55–75%, respectively, at a total organic aerosol loading,  $M_0$ , of  $15 \mu\text{g m}^{-3}$ . This is consistent with previous studies that have shown that SOA formation is less efficient under high- $\text{NO}_x$  conditions (Kroll et al., 2006; Presto and Donahue, 2006; Ng et al., 2007b), owing to formation of alkoxy radicals, followed by their decomposition. Large hydrocarbons, such as sesquiterpenes and long-chain alkanes, are exceptions to this behavior since isomerization of alkoxy radicals is more favorable (Ng et al., 2007a; Lim and Ziemann, 2005). Although the PAHs studied here also have relatively high MW, isomerization is not possible owing to the structure of the aromatic ring. As a result, fragmentation dominates under high- $\text{NO}_x$  conditions, leading to compounds that are more volatile.

The higher volatility of high- $\text{NO}_x$  SOA is evidenced by the dependence of SOA yields on  $M_0$ . Under high- $\text{NO}_x$ , aerosol growth was not observed until  $20 \mu\text{g m}^{-3}$  of the par-

1882

ent hydrocarbon had been reacted, indicating that gas-phase oxidation products are relatively volatile. Also, with the exception of 1,2-DMN, the final SOA yields increase with increasing  $M_0$ . This is characteristic of semivolatile partitioning, in which higher organic aerosol loading induces more partitioning into the aerosol phase. Under low- $\text{NO}_x$ , aerosol growth was observed almost immediately, as products have a much lower saturation concentration. The SOA yields are more or less constant, suggesting the SOA formed under low- $\text{NO}_x$  conditions is essentially nonvolatile in the range of  $M_0$  studied here.

The gas-phase PTR mass spectra obtained under high- and low- $\text{NO}_x$  photooxidation of naphthalene are shown in Fig. 3. These spectra are taken at the same fraction of naphthalene reacted, and thus the extents of reaction are similar. Figure 4 summarizes the proposed compounds for the major ion peaks. Under high- $\text{NO}_x$ , formylcinnamaldehyde, phthalaldehyde, and phthalic anhydride represent the largest peaks in the mass spectrum. We refer to these compounds as “ring-opening” products. Formation of these ring-opening compounds is consistent with the formation of alkoxy radicals via  $\text{RO}_2 + \text{NO}$  followed by decomposition of the alkoxy radicals. In a subsequent study, we will also show that phthalic acid, the hydrated form of phthalic anhydride, is the major product identified in the aerosol phase.

Under low- $\text{NO}_x$  conditions, these ring-opening products are formed in the absence of  $\text{NO}$ , as  $\text{RO}_2 + \text{RO}_2$  can also form alkoxy radicals, but the relative abundance is smaller than under high- $\text{NO}_x$  conditions. Reaction of  $\text{RO}_2$  with  $\text{HO}_2$  radicals produces hydroperoxides, but owing to the lack of structural information from the PTR-MS, one cannot distinguish between the protonated hydroperoxide ( $\text{C}_{10}\text{H}_{10}\text{O}_3 \cdot \text{H}^+$ ) and the water cluster of protonated formylcinnamaldehyde ( $\text{C}_{10}\text{H}_8\text{O}_2 \cdot \text{H}_2\text{O} \cdot \text{H}^+$ ). However, the relative abundance of other “ring-retaining” products, such as epoxides, naphthols, and quinones, is greater, because formation of the alkoxy radical is no longer dominant. Since these compounds and their oxidation products likely have a higher MW than the ring-opening products, their volatilities are lower, resulting in the overall low volatility of SOA formed and essentially constant yields under low- $\text{NO}_x$  conditions.

1883

### 4.3 Aerosol formation and yields

A widely-used semi-empirical model for SOA formation is based on the gas-particle equilibrium partitioning of semivolatile products (Odum et al., 1996, 1997), in which the SOA yield,  $Y$ , of a particular VOC, defined as the ratio of mass of SOA formed,  $\Delta M_0$ , to the mass of hydrocarbon reacted,  $\Delta \text{HC}$ , is given by

$$Y = \frac{\Delta M_0}{\Delta \text{HC}} = M_0 \sum_{i=1}^n \frac{\alpha_i K_{p,i}}{1 + K_{p,i} M_0} \quad (1)$$

where  $\alpha_i$  is the mass-based stoichiometric coefficient of semivolatile product  $i$ ,  $K_{p,i}$  is its gas-particle partitioning equilibrium constant, and  $M_0$  is the total mass concentration of the absorbing (aerosol) medium. In this study, since no organic aerosol seed was used, the total organic mass is equal to the mass of SOA formed ( $M_0 = \Delta M_0$ ). The final SOA yields for all experiments have been fitted to Eq. 1 for naphthalene, 1-MN and 2-MN photooxidation under high- $\text{NO}_x$  conditions, shown in Fig. 7. For naphthalene, a two-product model was used; for 1-MN and 2-MN, one product is sufficient to describe the experimental data. For 1,2-DMN under high- $\text{NO}_x$ , and all other PAHs under low- $\text{NO}_x$ , the SOA yields are constant. The fitting parameters are summarized in Table 3. The SOA yields for naphthalene and alkylnaphthalenes quantified in this study are on the order of 25–45% under high- $\text{NO}_x$  conditions, which are about three times those for light aromatics under similar photooxidation conditions (Ng et al., 2007b).

## 5 Estimation of SOA production from light aromatics, PAH and long-chain $n$ -alkanes

Using the compounds studied here as surrogates for PAHs, we estimate the contribution of PAH photooxidation to urban SOA formation from different anthropogenic sources.

1884

## 5.1 Diesel exhaust

Diesel exhaust is an important source of POA in the urban atmosphere (Schauer et al., 1996). As noted above, diesel exhaust, upon photooxidation, is also an SOA source (Robinson et al., 2007). The gas-phase component of diesel emissions consists primarily of light aromatics, PAHs and long-chain *n*-alkanes (Schauer et al., 1999a). For compound *i* emitted in the gas phase, the amount of SOA formed from photooxidation of that compound,  $\Delta M_{0,i}$ , expressed in units of  $\mu\text{g km}^{-1}$ , over a time  $\Delta t$ , can be estimated by

$$\Delta M_{0,i} = [HC_i](1 - e^{-k_{\text{OH},i}[\text{OH}]\Delta t}) \times Y_i \quad (2)$$

where  $HC_i$  is the amount of gas-phase emission of compound *i* in  $\mu\text{g km}^{-1}$ ,  $k_{\text{OH},i}$  is its reaction rate constant with OH radicals,  $[\text{OH}]$  is the concentration of OH radicals (taken to be  $2 \times 10^6 \text{ molec cm}^{-3}$ ) and  $Y_i$  is its SOA yield. To obtain a specific value of  $Y_i$ , we assume a total organic aerosol loading of  $15 \mu\text{g m}^{-3}$ . Emissions of each compound (and hence the amount of SOA formed), measured in  $\mu\text{g}$  per km driven, are taken from (Schauer et al., 1999a). OH reaction rate constants are taken from literature where available (Atkinson and Arey, 2003; Phoussongphouang and Arey, 2002), or estimated from structure activity relationships (Kwok and Atkinson, 1995) or extrapolation from similar compounds. The high- $\text{NO}_x$  SOA yields measured here are used in this analysis. For purpose of comparison, we use the SOA yields for benzene, toluene and *m*-xylene measured under similar photooxidation conditions (OH precursor, concentrations of  $\text{NO}_x$  and light intensity) (Ng et al., 2007b). SOA yields for *n*-alkanes were estimated based on modeling of gas-phase radical chemistry and gas-particle partitioning (Lim and Ziemann, 2005; Jordan et al., 2008).

Table 4 summarizes the contribution of each VOC identified to the total SOA formed from gas-phase photooxidation of these three classes of compounds. Although the gas-phase emissions are dominated by light aromatics, these compounds are estimated to account for only 14% of the SOA formed in first 3 h of photooxidation. This es-

1885

imate is consistent with laboratory results of photooxidation of diesel exhaust (Robinson et al., 2007), in which the “known” gas-phase precursors, consisting primarily of single-ring aromatics, account for at most 15% of the SOA formed. In addition to their relatively low SOA yields, the rate of oxidation of these compounds is slow, and hence only a small fraction is reacted over 3 h. On the contrary, although PAH emissions are a factor of 4 lower than those of light aromatics, their relatively fast reaction with OH radicals and high SOA yields lead to significant SOA production in the first 3 h, accounting for 4 times the amount formed from light aromatics. The contribution of PAH to SOA is still significant after 12 h of oxidation, at which point the SOA from PAH is about twice that from light aromatics.

Another important class of primary compounds is *n*-alkanes, which have been shown to have high aerosol yields under high- $\text{NO}_x$  conditions if the number of carbons is greater than 8 (Lim and Ziemann, 2005). In this calculation, SOA production from long-chain *n*-alkanes ranging from C12 to C20 is estimated to be higher than from light aromatics by a factor of 1.5 for the first 3 h. Considering only these three classes of compounds, PAHs are estimated to account for up to 54% of the SOA formed in the first 12 h of diesel exhaust photooxidation, representing a potentially large source of atmospheric SOA.

## 5.2 Wood burning

Wood burning has also been shown to produce a significant amount of SOA that cannot be accounted for by known volatile precursors (Grieshop et al., 2008b). Gas-phase emissions of aromatic compounds from fireplace wood burning are detailed in Schauer et al. (2001), in which 90% of the emissions (by mass) were identified. The gas-phase composition of wood combustion emissions depends on temperature, type of fuel, and method of burning (open vs. fireplace), and quantifying SOA from such sources can be difficult. Nonetheless, we repeat the previous analysis, and Table 5 summarizes the estimated SOA production from light aromatics and PAHs from fireplace burning of pine wood. Again, despite their relatively low emissions, the high SOA yields and

1886

rapid rate of oxidation of PAHs result in higher SOA formation. The amount of SOA from PAHs is estimated to be more than 4 times that from light aromatics after 12 h of oxidation, suggesting a potentially large source of SOA from oxidation of wood burning emissions.

### 5.3 Other anthropogenic sources

For gasoline exhaust, the emissions of PAHs relative to light aromatics are low such that their contribution to SOA formation is likely insignificant (Schauer et al., 2002a). For other sources, such as meat and vegetable cooking, the emissions of light aromatics and PAHs are insignificant (Schauer et al., 1999b, 2002b). One class of compounds not addressed in this analysis is  $n$ -C<sub>10</sub>–C<sub>14</sub> aldehydes, which are emitted in significant quantities from each of the sources considered here. To the authors' knowledge, SOA yields for these compounds have not been quantified and could potentially contribute to urban SOA formation.

We stress that the analysis presented here constitutes a first-order estimate of SOA production from PAHs. The calculations are based on emissions measured from single sources, rather than on overall emissions inventories. Future work will address the prediction of SOA formation based on atmospheric emissions inventories of gas-phase PAHs.

## 6 Implications

In this study, we have investigated the formation of SOA from gas-phase photooxidation of four 2-ring PAHs (naphthalene, 1-MN, 2-MN and 1,2-DMN) under high- and low-NO<sub>x</sub> conditions. Under high-NO<sub>x</sub> and at organic aerosol loadings between 10 and 40 μg m<sup>-3</sup>, SOA yields ranged from 0.19 to 0.30 for naphthalene, 0.19 to 0.39 for 1-MN and 0.26 to 0.45 for 2-MN. The SOA products were found to be semivolatile, consistent with the gas-phase products being dominated by ring-opening compounds. Under low-

1887

NO<sub>x</sub> conditions, the SOA yield was found to be constant in the range of  $M_0$  studied, and was measured to be 0.73, 0.68, and 0.58, for naphthalene, 1-MN, and 2-MN, respectively. More ring-retaining products were found in the gas phase under low-NO<sub>x</sub> conditions, contributing to the low volatility of the SOA and constant SOA yields. The SOA formed from high-NO<sub>x</sub> photooxidation of 1,2-DMN was also nonvolatile under these loadings.

When applying these measured yields to a simple model of SOA formation from primary emissions, SOA from PAH oxidation is found to contribute a significant amount. Traditionally thought to be the most important anthropogenic SOA precursors, light aromatics are estimated to produce 3–5 times less SOA than from PAHs, based on the emissions inventories employed here. PAHs represent a potentially large source of urban SOA, accounting for up to 54% of SOA formed from diesel exhaust, and up to 80% of SOA from fireplace wood burning. Long-chain  $n$ -alkanes are also expected to produce a smaller but significant fraction of SOA from diesel exhaust.

To properly account for SOA formation from anthropogenic sources, photooxidation of low-volatility precursors, such as PAH and  $n$ -alkanes, should be included in atmospheric SOA models, likely increasing the amount of predicted anthropogenic SOA significantly. Other IVOCs, such as aliphatic aldehydes and large olefins, could also be an important source of SOA, but further study is required to evaluate their potential to form SOA. In a forthcoming study, we will examine the mechanism of PAH oxidation and the composition of the SOA generated from the experiments reported here to assess the atmospheric relevance of these results.

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1888



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1891

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1892

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1893

**Table 1.** Experimental conditions and results.

Date (2008)	Compound	[HC] <sub>0</sub> , ppb	[NO] <sub>0</sub> , ppb	[NO <sub>2</sub> ] <sub>0</sub> , ppb	V <sub>0</sub> <sup>a</sup> , μm <sup>3</sup> cm <sup>-3</sup>	ΔHC, μg m <sup>-3</sup>	ΔM <sub>0</sub> , μg m <sup>-3</sup>	Yield
6/12	naphthalene	7.5	<sup>b</sup>	<sup>c</sup>	12.9	39.5±2.5	29.0±2.6	0.73±0.11
6/14	naphthalene	4.4	<sup>b</sup>	<sup>c</sup>	14.2	22.8±2.1	17.0±1.7	0.74±0.15
6/17	naphthalene	6.0	<sup>b</sup>	<sup>c</sup>	13.9	31.2±2.1	22.9±2.2	0.73±0.12
6/19	naphthalene	2.7	<sup>b</sup>	<sup>c</sup>	14.5	14.4±2.0	10.2±1.5	0.71±0.21
6/13	naphthalene	12.1	411	83	14.5	58.6±2.0	10.9±1.5	0.19±0.03
6/16	naphthalene	21.2	304	73	13.6	92.7±2.1	27.7±3.0	0.30±0.04
6/18	naphthalene	9.5	351	73	15.2	39.6±2.0	8.0±1.2	0.20±0.04
6/20	naphthalene	17.0	411	66	15.4	65.0±1.6	13.6±1.7	0.21±0.03
						83.9±2.1 <sup>f</sup>	22.8±2.2 <sup>f</sup>	0.27±0.03 <sup>f</sup>
7/16	2-MN	7.5	<sup>b</sup>	<sup>c</sup>	13.2	58.6±3.4	35.0±3.1	0.60±0.09
7/22	2-MN	4.4	<sup>b</sup>	<sup>c</sup>	14.6	31.5±3.0	18.6±1.8	0.59±0.11
7/24	2-MN	6.0	<sup>b</sup>	<sup>c</sup>	12.9	22.5±3.3	13.1±1.5	0.58±0.15
7/27	2-MN	2.7	<sup>b</sup>	<sup>c</sup>	13.3	43.3±3.0	24.6±2.1	0.57±0.09
7/21	2-MN	21.2	398	52	14.1	28.0±3.4	7.4±0.9	0.26±0.06
7/23	2-MN	9.5	376	51	14.2	47.3±3.5	18.8±1.9	0.40±0.07
7/25	2-MN	17.0	<sup>d</sup>	55	12.9	70.7±3.4	31.6±2.6	0.45±0.06
7/29	2-MN	7.5	<sup>d</sup>	61	13.4	60.0±3.6	24.3±2.2	0.40±0.06
8/2	1-MN	7.3	<sup>b</sup>	<sup>c</sup>	11.4	42.9±3.9	30.5±2.5	0.71±0.12
8/5	1-MN	9.9	<sup>b</sup>	<sup>c</sup>	13.1	57.4±3.4	37.6±3.0	0.66±0.09
8/8	1-MN	2.8	<sup>b</sup>	<sup>c</sup>	13.8	15.7±3.4	10.7±1.1	0.68±0.22
8/17	1-MN	5.9	<sup>b</sup>	<sup>c</sup>	14.0	35.2±3.6	24.0±2.4	0.68±0.14
8/4	1-MN	10.5	<sup>d</sup>	77	13.6	51.8±3.0	16.5±1.6	0.32±0.05
8/7	1-MN	5.3	333	55	13.5	30.1±3.4	5.6±0.9	0.19±0.05
8/9	1-MN	14.6	388	62	14.7	79.6±3.6	30.9±2.6	0.39±0.05
8/11	1-MN	16.8	401	56	14.6	91.0±3.6	34.9±3.1	0.38±0.05
9/10	1,2-DMN	34.8	392	51	14.7	215.7±6.8	66.9±5.3	0.31±0.03
9/12	1,2-DMN	12.3	397	60	14.9	76.8±4.4	23.1±2.3	0.30±0.05
9/14	1,2-DMN	4.1	377	55	14.2	25.4±3.3	7.8±1.4	0.31±0.10

<sup>a</sup> V<sub>0</sub>: volume concentration of ammonium sulfate seed

<sup>b</sup> below detection limit of 2 ppb

<sup>c</sup> NO<sub>2</sub> concentration was not measured by gas-chromatograph in these experiments due to H<sub>2</sub>O<sub>2</sub> interference with NO<sub>2</sub> signal, but inferred to be <2 ppb from NO/NO<sub>x</sub> analyzer.

<sup>d</sup> NO concentration not measured, but expected to be similar to that of other high-NO<sub>x</sub> experiments.

<sup>e</sup> Nucleation experiments not shown; with the exception of seed concentration, conditions were similar to those of experiments shown here.

<sup>f</sup> After second addition of HONO

1894

**Table 2.** Effective SOA densities ( $\text{g cm}^{-3}$ ), obtained in nucleation experiments.

	high- $\text{NO}_x$	low- $\text{NO}_x$
naphthalene	1.48	1.55
1-MN	1.4	1.43
2-MN	1.35	1.3
1,2-DMN	<sup>a</sup>	—

<sup>a</sup> Not measured; assumed to be 1.4

1895

**Table 3.** SOA yield parameters.

	$\alpha_1$	$K_1, \text{m}^3 \mu\text{g}^{-1}$	$\alpha_2$	$K_2, \text{m}^3 \mu\text{g}^{-1}$	Yield at $M_0 = 15 \mu\text{g m}^{-3}$
<b>high-<math>\text{NO}_x</math></b>					
naphthalene	0.21	0.59	1.07	0.0037	0.26
1-MN	0.50	0.11	—	—	0.33
2-MN	0.55	0.13	—	—	0.38
1,2-DMN	0.31	<sup>a</sup>	—	—	0.31
<b>low-<math>\text{NO}_x</math></b>					
naphthalene	0.73	<sup>a</sup>	—	—	0.73
1-MN	0.68	<sup>a</sup>	—	—	0.68
2-MN	0.58	<sup>a</sup>	—	—	0.58

<sup>a</sup> Constant yield

1896

**Table 4.** Estimated SOA production from gas-phase oxidation of diesel exhaust (Schauer et al., 1999a).

	$[HC_i]$ , $\mu\text{g km}^{-1}$	$k_{\text{OH}}$ , $\text{cm}^3 \text{ molec}^{-1} \text{ s}^{-1}$	$Y$	$\Delta t=3 \text{ h}$ $\Delta M_{0,i}$ , $\mu\text{g km}^{-1}$	$\Delta t=12 \text{ h}$ $\Delta M_{0,i}$ , $\mu\text{g km}^{-1}$
<b>light aromatics</b>					
benzene	2740	$1.22 \times 10^{-12} \text{ a}$	$0.19^{\text{f}}$	14	53
toluene	3980	$5.63 \times 10^{-12} \text{ a}$	$0.10^{\text{f}}$	47	158
ethylbenzene	470	$7.00 \times 10^{-12} \text{ a}$	$0.10^{\text{g}}$	7	22
xylene	2330	$1.36 \times 10^{-11} \text{ a}$	$0.06^{\text{h}}$	35	96
<i>o</i> -xylene	830	$2.31 \times 10^{-11} \text{ a}$	$0.06^{\text{h}}$	19	43
<i>n</i> -propylbenzene	100	$1.43 \times 10^{-11} \text{ a}$	$0.10^{\text{g}}$	3	7
<i>p</i> -ethyltoluene	520	$1.18 \times 10^{-11} \text{ a}$	$0.10^{\text{g}}$	12	34
<i>m</i> -ethyltoluene	210	$1.86 \times 10^{-11} \text{ a}$	$0.10^{\text{g}}$	7	17
1,3,5-TMB	260	$5.67 \times 10^{-11} \text{ a}$	$0.06^{\text{h}}$	11	15
1,2,4-TMB	880	$3.25 \times 10^{-11} \text{ a}$	$0.06^{\text{h}}$	27	49
<b>Total</b>				<b>182</b>	<b>496</b>
<b>PAHs</b>					
naphthalene	617	$2.30 \times 10^{-11} \text{ a}$	0.26	62	136
2-methylnaphthalene	611	$4.86 \times 10^{-11} \text{ b}$	0.38	152	231
1-methylnaphthalene	378	$4.09 \times 10^{-11} \text{ b}$	0.33	73	120
C2-naphthalenes	542	$6.00 \times 10^{-11} \text{ c}$	0.31	122	167
C3-naphthalenes	240	$8.00 \times 10^{-11} \text{ d}$	0.31	61	74
C4-naphthalenes	97.3	$8.00 \times 10^{-11} \text{ d}$	0.31	25	30
other PAHs	886.7	$8.00 \times 10^{-11} \text{ d}$	0.31	226	275
<b>Total</b>				<b>721</b>	<b>1033</b>
<b><i>n</i>-alkanes</b>					
dodecane	503	$1.32 \times 10^{-11} \text{ a}$	$0.02^{\text{j}}$	2	5
tridecane	477	$1.51 \times 10^{-11} \text{ a}$	$0.03^{\text{j}}$	4	10
tetradecane	629	$1.79 \times 10^{-11} \text{ a}$	$0.05^{\text{j}}$	10	23
pentadecane	398	$2.07 \times 10^{-11} \text{ a}$	$0.08^{\text{j}}$	11	25
hexadecane	711	$2.32 \times 10^{-11} \text{ a}$	$0.12^{\text{j}}$	34	75
heptadecane	614	$2.85 \times 10^{-11} \text{ e}$	$0.20^{\text{j}}$	55	110
octadecane	601	$3.51 \times 10^{-11} \text{ e}$	$0.30^{\text{j}}$	96	172
nonadecane	411	$4.32 \times 10^{-11} \text{ e}$	$0.42^{\text{j}}$	105	168
eicosane	271.8	$5.31 \times 10^{-11} \text{ e}$	$0.56^{\text{j}}$	104	151
<b>Total</b>				<b>420</b>	<b>740</b>

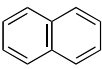
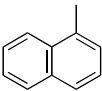
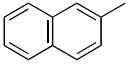
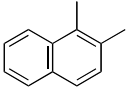
<sup>a</sup> Atkinson and Arey (2003); <sup>b</sup> Phousongphouang and Arey (2002); <sup>c</sup> Average of all dimethylnaphthalenes in Phousongphouang and Arey (2002); <sup>d</sup> Extrapolated from dimethylnaphthalenes, assumed to be same for all other PAHs; <sup>e</sup> Estimated from structure-reactivity relationships (Kwok and Atkinson, 1995); <sup>f</sup> Ng et al. (2007b); <sup>g</sup> Assumed to be the same as toluene; <sup>h</sup> Assumed to be the same as *m*-xylene; <sup>i</sup> Estimated from modeling of gas-phase chemistry and gas/particle partitioning (Jordan et al., 2008); <sup>j</sup> Extrapolated from smaller *n*-alkanes.

1897

**Table 5.** Estimated SOA production from gas-phase oxidation of emissions from fireplace burning of pine wood (Schauer et al., 2001).

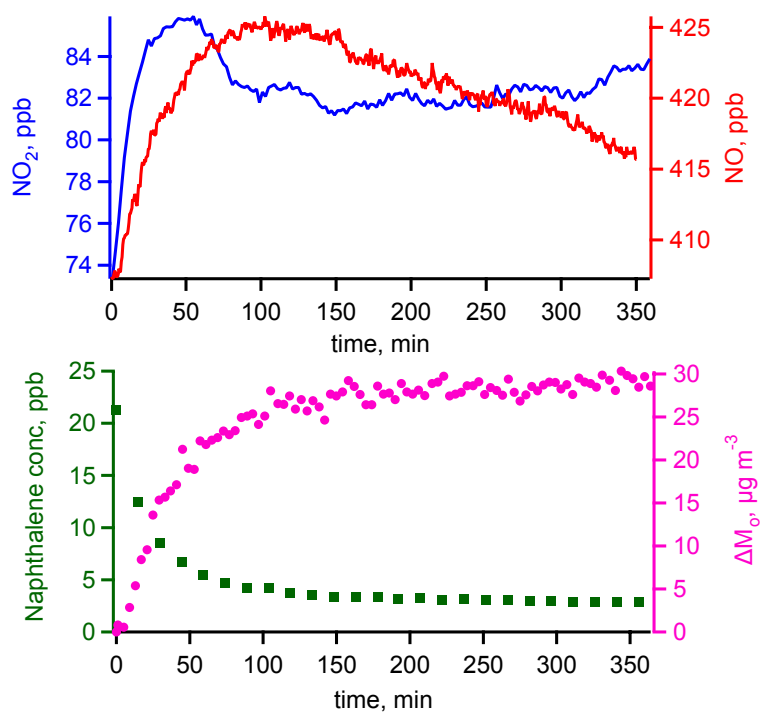
	$[HC_i]$ , $\text{mg kg}^{-1}$	$k_{\text{OH}}$ , $\text{cm}^3 \text{ molec}^{-1} \text{ s}^{-1}$	$Y$	$\Delta t=12 \text{ h}$ $\Delta M_{0,i}$ , $\text{mg kg}^{-1}$
<b>light aromatics</b>				
benzene	383	$1.22 \times 10^{-12}$	0.19	7.5
toluene	158	$5.63 \times 10^{-12}$	0.10	6.3
ethylbenzene	22.9	$7.00 \times 10^{-12}$	0.10	1.1
xylene	60	$1.36 \times 10^{-11}$	0.06	2.5
<i>o</i> -xylene	18.1	$2.31 \times 10^{-11}$	0.06	0.9
<b>Total</b>				<b>18.2</b>
<b>PAHs</b>				
naphthalene	227	$2.30 \times 10^{-11}$	0.26	50.0
2-methylnaphthalene	15	$4.86 \times 10^{-11}$	0.38	5.7
1-methylnaphthalene	10.6	$4.09 \times 10^{-11}$	0.33	3.4
C2-naphthalenes	13.8	$6.00 \times 10^{-11}$	0.31	4.3
other PAHs	57.0	$8.00 \times 10^{-11}$	0.31	17.7
<b>Total</b>				<b>80.9</b>

1898

Compound	Structure	OH rate constant, $\text{cm}^3 \text{molec}^{-1} \text{s}^{-1}$
naphthalene		$2.44 \times 10^{-11}$
1-methylnaphthalene (1-MN)		$4.09 \times 10^{-11}$
2-methylnaphthalene (2-MN)		$4.86 \times 10^{-11}$
1,2-dimethylnaphthalene (1,2-DMN)		$5.96 \times 10^{-11}$

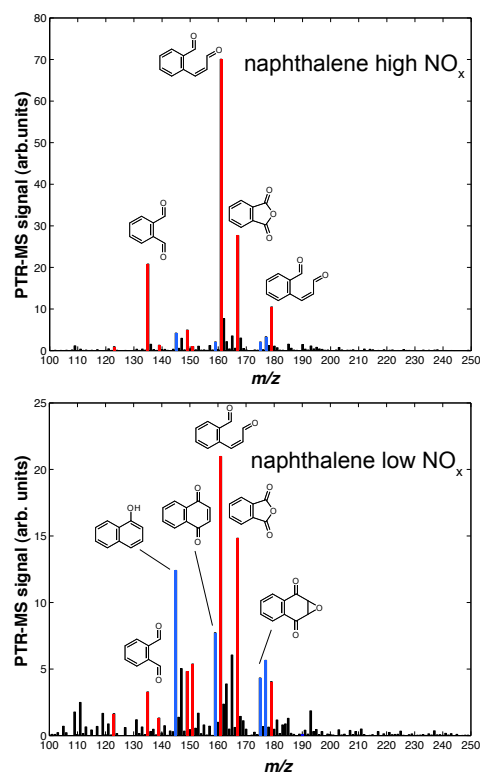
**Fig. 1.** Structure of PAHs studied and OH rate constants (Phousongphouang and Arey, 2002).

1899



**Fig. 2.** Reaction profile during a typical high- $\text{NO}_x$  experiment. The mass concentration of SOA has been corrected for particle wall loss.

1900



**Fig. 3.** PTR-MS spectra under high- and low- $\text{NO}_x$  conditions. The fractions of naphthalene reacted are 0.66 and 0.67, respectively. The red peaks indicate ions associated with ring-opening products, and the blue peaks indicate those associated with ring-retaining products. The proposed structures for these ion peaks are given in Fig. 4.

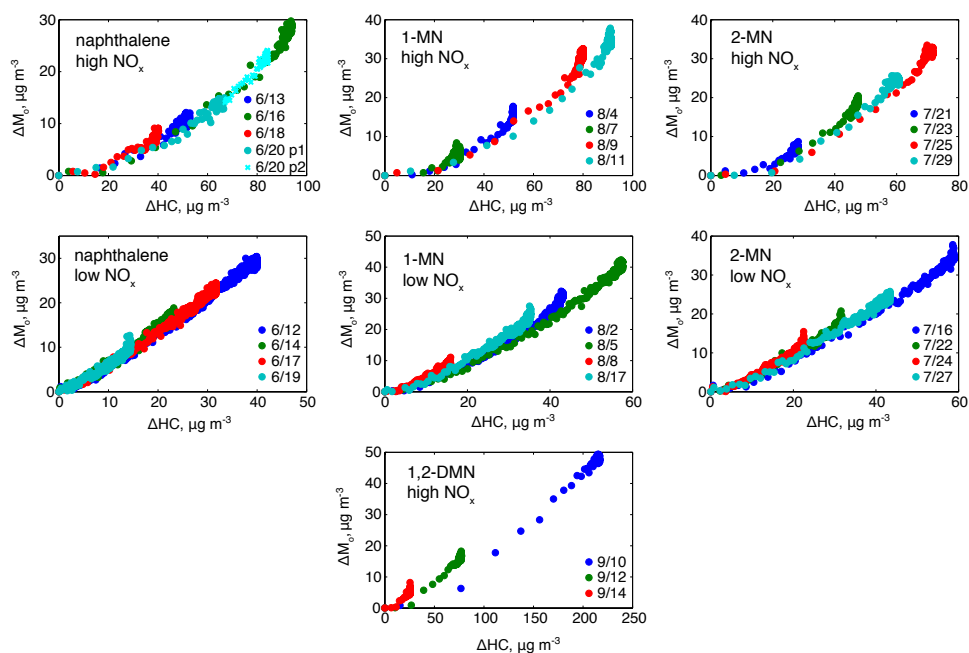
1901

Name	Proposed Structure	PTR-MS $m/z$ (+)	Name	Proposed Structure	PTR-MS $m/z$ (+)
<b>Ring-opening products</b>			<b>Ring-retaining products</b>		
benzoic acid		123	naphthol		145
phthalaldehyde		135	1,4-naphthoquinone		159
hydroxybenzoic acid		139	2,3-epoxy-1,4-quinone		175
phthalic anhydride		149, 167	4-hydroxy-2,3-epoxy-carbonyl		177*
formylbenzoic acid		151			
2-formylcinnamaldehyde		161, 179			
formylcinnamic acid		177*			

\* 2 proposed isomers for  $m/z$  177

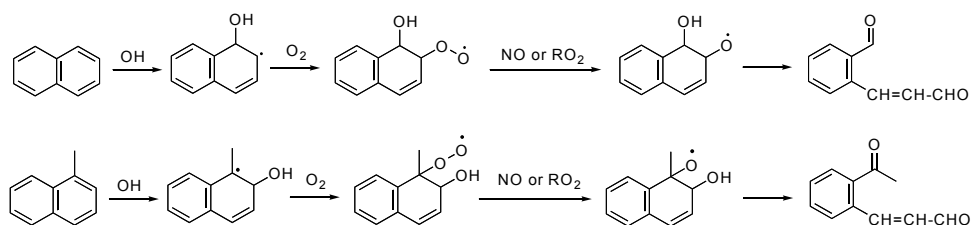
**Fig. 4.** Proposed compounds and structures for ion peaks observed by PTR-MS. The ring-opening gas-phase products are marked red and the gas-phase ring-retaining products are marked blue, to correspond to Fig. 3.

1902



**Fig. 5.** Time-dependent SOA growth curves for all experiments. In one naphthalene high- $\text{NO}_x$  experiment (6/20), more HONO was added after the first injection was consumed. SOA growth before the second HONO addition is marked light blue and labelled “6/20 p1”, and the growth after the second addition is marked bright cyan and labelled “6/20 p2”.

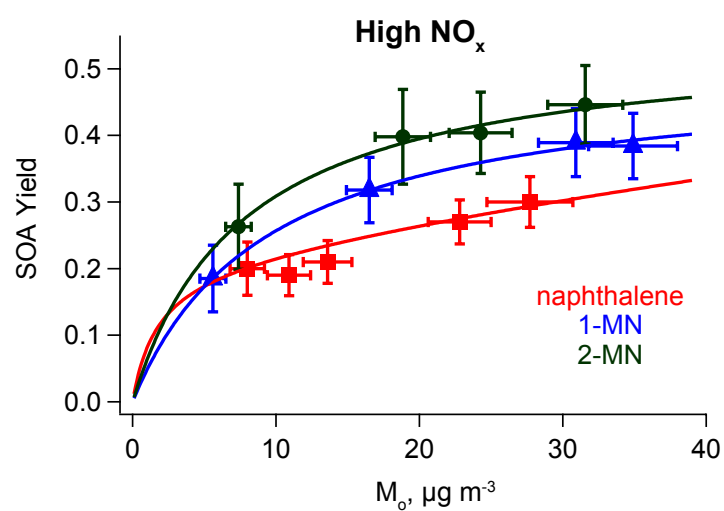
1903



**Fig. 6.** Mechanism for oxidation of naphthalene and 1-MN to form first-generation dicarbonyl fragmentation product (Wang et al., 2007).

1904





**Fig. 7.** SOA yield as a function of organic loading for naphthalene, 1-MN and 2-MN under high-NO<sub>x</sub> conditions. The fitting parameters used are summarized in Table 3. Under low-NO<sub>x</sub> conditions, the SOA yields are essentially constant.