

Formation of secondary organic aerosol and oligomers from the ozonolysis of enol ethers

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Abstract

Formation of secondary organic aerosol has been observed in the gas phase ozonolysis of a series of enol ethers, among them several alkyl vinyl ethers (AVE, $\text{ROCH}=\text{CH}_2$), such as ethyl, propyl, *n*-butyl, *iso*-butyl, *t*-butyl vinyl ether, and ethyl propenyl ether (EPE, $\text{C}_2\text{H}_5\text{OCH}=\text{CHCH}_3$). The ozonolysis has been studied in a 570 l spherical glass reactor at atmospheric pressure (730 Torr) and temperature (296 K). Gas phase reaction products were investigated by in-situ FTIR spectroscopy, and secondary organic aerosol (SOA) formation was monitored by a scanning mobility particle sizer (SMPS). The chemical composition of the formed SOA was analysed by a hybrid mass spectrometer using electrospray ionization (ESI). The main stable gas phase reaction product is the respective alkyl formate $\text{ROC}(\text{O})\text{H}$, formed with yields of 60 to 80%, implying that similar yields of the corresponding Criegee Intermediates (CI) CH_2O_2 for the AVE and CH_3CHO_2 for EPE are generated. Measured SOA yields are between 2 to 4% for all enol ethers. Furthermore, SOA formation is strongly reduced or suppressed by the presence of an excess of formic acid, which acts as an efficient CI scavenger.

Chemical analysis of the formed SOA by ESI(+)/MS-TOF allows to identify oligomeric compounds in the mass range 200 to 800 u as its major constituents. Repetitive chain units are identified as CH_2O_2 (mass 46) for the AVE and $\text{C}_2\text{H}_4\text{O}_2$ (mass 60) for EPE and thus have the same chemical compositions as the respective major Criegee Intermediates formed during ozonolysis of these ethers. The oligomeric structure and chain unit identity are confirmed by HPLC/ESI(+)/MS-TOF and ESI(+)/MS/MS-TOF experiments, whereby successive and systematic loss of a fragment with mass 46 for the AVE (and mass 60 for EPE) is observed. It is proposed that the oligomer has the following basic structure of an oligoperoxide, $[\text{CH}(\text{R})\text{-O-O}]_n^-$, where $\text{R}=\text{H}$ for the AVE and $\text{R}=\text{CH}_3$ for the EPE. Oligoperoxide formation is thus suggested to be another, so far unknown reaction of stabilized Criegee Intermediates in the gas phase ozonolysis of oxygen-containing alkenes leading to SOA formation.

1 Introduction

Alkyl vinyl ethers are increasingly produced as solvents to be used in the chemical industry. Because of their reasonable volatility this class of compounds will result in a significant emission into the urban atmosphere. Consequently alkyl vinyl ethers will be oxidized by OH- and NO₃-radicals and O₃.

A few investigations have recently been reported on the kinetics and product distribution of simple alkyl vinyl ethers. Thiault et al. (2002) studied the OH and O₃ initiated oxidation of ethyl vinyl ether (EVE, C₂H₅O-CH=CH₂). The main products from both reactions were formaldehyde and ethyl formate C₂H₅OC(O)H. Additional studies performed in the EUPHORE chamber revealed the formation of secondary organic aerosols (SOA) (George et al., 2005).

Klotz et al. (2004) performed a product study of the gas-phase reactions of O₃, OH and NO₃ with methyl vinyl ether (MVE, CH₃O-CH=CH₂). The main products from all reactions were formaldehyde and methyl formate CH₃OC(O)H; in the case of the O₃ reaction, hydroperoxymethyl formate, methanol, CO₂ and CO were also produced. Interestingly, particle formation was observed in the ozonolysis, but not in the OH or NO₃ radical reactions. The observed aerosol profiles showed typical behaviour associated with homogeneous nucleation. The mean size of the particles measured at the end of the experiment (1 to 3 h) was typically 60 nm in diameter. Methyl vinyl ether with only three carbon atoms, is the smallest compound studied that produces aerosol during its atmospheric degradation. Usually, seven or more C atoms are required from non-cyclic hydrocarbons, six from cyclic species to initiate SOA formation (Seinfeld and Pankow, 2003; Seinfeld and Pandis, 1998).

A detailed investigation of the ozonolysis of six unsaturated alkyl ethers has been performed in a laboratory reaction chamber: ethyl vinyl ether (EVE), ethyl propenyl ether (EPE), propyl vinyl ether (PVE), *n*-butyl vinyl ether (BVE), *t*-butyl vinyl ether (TBVE), and *iso*-butyl vinyl ether (IBVE). Gas phase products have been quantified by FTIR spectroscopy and SOA have been observed by a SMPS system and char-

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acterized by an HPLC coupled to a hybrid mass spectrometer (Sadezky, 2005). This paper reports the formation and identification of a series of oligomeric products in the ozonolysis of all the studied unsaturated alkyl ethers.

2 Experimental

Experiments in the laboratory were performed in a 570-liter spherical glass reactor at room temperature in synthetic air at a total pressure of 730 Torr. A detailed description of this setup has been described in earlier publications (Neeb et al., 1998; Winterhalter et al., 2000). Ozone was produced by a mercury pen-ray lamp inside the reactor, prior to the addition of the mixture of vinyl ether and synthetic air (reaction start). The concentrations of reactants and reaction products were followed by Fourier Transform infrared spectroscopy (FTIR). The aerosol concentration and size distribution was monitored with a scanning mobility particle sizer (SMPS, TSI 3936). The SMPS consists of an electrostatic classifier (TSI 3080) with a long differential mobility analyzer, (LDMA; TSI 3081) and an ultrafine condensation particle counter (CPC; TSI 3025A) as detector.

Experiments were performed with initial ether mixing ratios varying between 200 ppb and 9 ppm, and similar or slightly lower ozone mixing ratios. For the studies involving observation of SOA formation profiles, mixing ratios between 200 and 400 ppb of ether and ozone were used, whereas for experiments with chemical analysis by mass spectrometry of collected aerosols, 8 ppm of ozone and 9 ppm of ether were used. In order to prevent reactions of vinyl ethers with OH radicals, which are known to be generated during the ozonolysis of alkenes (Finlayson et al., 1972), cyclohexane (excess, 300 ppm) was added in some experiments. In another type of experiments, excess HCOOH (20 ppm) was added as scavenger for stabilized Criegee Intermediates. All chemicals were commercially available (purity >95%) and used without further purification.

The aerosol formed in the laboratory experiments was collected during 20 to 25 min

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on Teflon (PTFE) filters (45 mm diameter, 0.45 μm pore size), using a flow rate of 14 L min^{-1} . After collection the filters were extracted in a 7 cm^3 glass flask with 3 ml pure methanol (HPLC grade), and stored at -20°C until analysis.

For chromatographic separation, 10 μL of this extract were directly injected into the HPLC system, which consisted of a thermostated autosampler (Series 200, Perkin Elmer), a degasser and a quaternary pump (both 1100 Series, Agilent Technologies). The analytical column (250 mm \times 2 mm I.D., 5 μm particle size) contains a ReproSil-Pur C₁₈-AQ stationary phase (Dr. Maisch GmbH). The eluents were 0.1% formic acid in deionized water (eluent A) and acetonitrile (HPLC-grade) (eluent B). The gradient of the mobile phase was as follows: 0% B for 0.5 min, gradient to 15% B in 2.5 min, gradient to 100% B in 16 min, isocratic for 6.5 min. The flow of the mobile phase was 400 $\mu\text{L}/\text{min}$.

Chemical constituents were detected by a hybrid mass spectrometer (quadrupole and time-of-flight) QSTAR (Applied Biosystems MDS SCIEX) with an electrospray ion source. The extraction solution was either directly injected (10 $\mu\text{L}/\text{min}$) or previously separated by HPLC as described before. The electrospray ion source (Turbolonspray) was operated in the positive mode at 400°C and an ionization voltage of +3.4 kV. The declustering potential was +30 V, and the focussing potential (focus ring) was +100 V. For tandem Q-TOF experiments, the collision energy was between 15 and 30 eV with CAD (collision gas) set to 2. Instrument control, spectra treatment and calculations of elemental compositions were done with the software Analyst (Applied Biosystems MDS SCIEX). Further instrumental and analytical details are reported by Römpf (2003).

3 Results

3.1 General aspects and gas phase mechanism of the ozonolysis of enol ethers

The general mechanism of the ozonolysis of unsaturated ethers is displayed in Fig. 1. The initial product formed is the primary ozonide (1,2,3-trioxolane), which is unstable

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and decomposes into a carbonyl oxide, called the Criegee Intermediate (CI), and a primary carbonyl compound. The ozonolysis of the ether molecules investigated in this study produces CIs of the type CH_2OO (C₁-CI, $\text{R}'=\text{H}$) for the alkyl vinyl ethers (AVE), and CH_3CHOO (C₂-CI, $\text{R}'=\text{CH}_3$) for ethyl propenyl ether (EPE), together with RO-CHOO (alkoxy-substituted CI). The R group consists of C_2H_5 for EVE and EPE, C_3H_7 for PVE and C_4H_9 for the three different butyl vinyl ethers. The corresponding primary carbonyl compounds consist of an alkyl formate and formaldehyde or acetaldehyde. The CIs formed from the decomposition of the primary ozonide are formed in excited states, which then either decompose into various products or become collisionally stabilized. Nearly half of the excited CIs are stabilized, and it has been established that stabilized CIs (SCIs) undergo in the gas phase a variety of bimolecular reactions with carbonyl compounds, HCOOH and H_2O to form secondary ozonides (1,2,4-trioxolanes), hydroperoxy formates and hydroxyhydroperoxides, respectively (Calvert et al., 2000).

The yields of the different gas phase products have been determined by FTIR spectroscopy for EVE and EPE, from which the branching ratio of the splitting of the primary ozonide into both pathways was deduced to be $(71 \pm 13)\%$ for the "ethyl formate + C₁-CI" channel for EVE, and $(83 \pm 13)\%$ for the "ethyl formate + C₂-CI" channel for EPE (Sadezky, 2005). It has to be stressed that the yield of the C₁ and C₂ SCIs is estimated to be nearly 20–40% per reacted enol ether (Sadezky, 2005).

The gas-phase chemistry of the ozonolysis of the unsaturated ethers will not be treated in this paper, but will be published elsewhere (Sadezky et al., 2006¹). An elaborate mechanism of the gas phase ozonolysis of methyl vinyl ether has been presented by Klotz et al. (2004), and their results show the same trend as observed in our studies (Sadezky, 2005).

Besides the classical mechanism of the ozonolysis of unsaturated ethers in the gas phase, as presented above, there exists another mechanism in the liquid phase. Keul et al. (1985) investigated the ozonolysis of different alkyl vinyl ethers in a variety of

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solvents at -78°C , and observed major yields of alkoxy-1,2-dioxolanes, besides the expected alkoxy-1,2,4-trioxolanes. These products arose from the addition of carbonyl oxide CH_2OO , produced upon ozonolysis of vinyl ethers from decomposition of the primary ozonide, to the double bond of the vinyl ethers by 1,3-cycloaddition (Fig. 2).

3.2 Aerosol formation from gas-phase ozonolysis of AVE

3.2.1 Size distribution of SOA particles

Typical time evolution of the particle number size distribution during experiments of various alkyl vinyl ethers (AVE) is presented in Fig. 3 for ethyl vinyl ether (EVE) (initial mixing ratios 200 to 400 ppb of reactants EVE and O_3). The SOA appears as a single aerosol mode, whose time evolution is sequentially displayed with different colours for different reaction times of the EVE ozonolysis. Formation of detectable particles ($>10\text{ nm}$) is observed after 3 min, given by the turquoise-coloured size distribution. A maximum total particle concentration of 5×10^5 to 1×10^6 particle/ cm^3 is reached after 5 min at mode diameters (top point of the aerosol mode) around 20 nm (pink-coloured size distribution). With increasing reaction time, the total particle concentration decreases due to particle coagulation, while the particle diameter increases due to condensation of gaseous reaction products onto the particle surfaces and to coagulation. At the end of the reaction (after 45 min), the SOA size distribution is stable at mode diameters around 50 nm and a total particle concentration of 1×10^5 particle/ cm^3 . Similar observations were made for the other AVE under similar experimental conditions and initial reactant mixing ratios.

As no seed aerosol is used, and particle background concentrations ($>3\text{ nm}$) are measured to be lower than 200 particle/ cm^3 , SOA particle formation necessarily occurs by homogeneous nucleation of low-volatile reaction products from gas phase ozonolysis of the enol ethers.

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3.2.2 SOA yields and partitioning coefficients

Total aerosol mass concentrations ($\mu\text{g}/\text{m}^3$) were calculated from the measured number size distributions by assuming spherical particles with a density of $1\text{ g}/\text{cm}^3$. Figure 4 shows for propyl vinyl ether (PVE) plots of the total mass concentration M_0 of aerosol formed as a function of reaction time (Fig. 4a) and of the mass concentration of the ether reacted, ΔPVE (Fig. 4b). The linear part of the plot in Fig. 4b characterizes the SOA formation before a final and stable total SOA mass concentration $M_0 = 18\text{ }\mu\text{g}/\text{m}^3$ for PVE is reached. These final SOA mass concentrations for the other ethers, which are reached after 15 to 20 min, were: $M_0 = 13\text{ }\mu\text{g}/\text{m}^3$, $25\text{ }\mu\text{g}/\text{m}^3$, $14\text{ }\mu\text{g}/\text{m}^3$, and $7\text{ }\mu\text{g}/\text{m}^3$ for ethyl vinyl ether (EVE), *n*-butyl vinyl ether (BVE), *iso*-butyl vinyl ether (IBVE) and *t*-butyl vinyl ether (TBVE), respectively. It is clearly evident from Fig. 4b that a certain quantity of the AVE must have reacted before SOA is formed and observed. The concentration of enol ether reacted until formation of *detectable* aerosol particles is observed is referred to as threshold quantity $\Delta\text{AVE}_{\text{thresh}}$. $\Delta\text{AVE}_{\text{thresh}}$ corresponds to the intersection of the linear section with the x-axis representing the mass of reacted AVE, ΔAVE . The slope of this linear section (Fig. 4b) gives the aerosol yield, a , hereby ignoring the threshold quantity $\Delta\text{AVE}_{\text{thresh}}$ reacted before:

$$M_0 = a \times \Delta\text{AVE} + b \quad (1)$$

with M_0 = total aerosol mass concentration ($\mu\text{g}/\text{m}^3$), ΔAVE = mass of reacted AVE ($\mu\text{g}/\text{m}^3$), and a = aerosol yield.

The aerosol profiles can also be evaluated by the adsorption/partitioning model developed by Odum et al. (1996). In this model hypothetical semi-volatile compounds i are partitioned between the gas and aerosol phase, as expressed in the Eq. (2):

$$Y_f = M_0 \sum \alpha_i K_i / (1 + K_i M_0) \quad (2)$$

Y_f is the fractional SOA yield calculated by

$$Y_f = M_0 / \Delta\text{AVE} \quad (3)$$

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The partitioning coefficients K_i and the formation yields α_i for the semi-volatile compounds i can be obtained from a least-square fit of the observed aerosol yield Y_f as a function of the aerosol mass M_0 . Here it is assumed that only one semi-volatile compound i is partitioned for every AVE, leading to Eq. (4):

$$Y_f = M_0[\alpha_i K_i / (1 + K_i M_0)] \quad (4)$$

The corresponding plots and least-square fits for the different AVEs are shown in Fig. 5, combining the results of three to four repetitive ozonolysis experiments for each AVE under similar experimental conditions.

Table 1 summarizes the results of both methods for the evaluation of SOA formation profiles in the reaction of the five studied AVE with ozone. The SOA mass yields a are in very good agreement with the formation yields α_i of the hypothetical semi-volatile compound i . For EVE, PVE, BVE and IBVE typical yields range from 3.5 to 4.7%. Threshold quantities ΔAVE_{thresh} vary between 37 and 58 ppb. For TBVE the aerosol yield is much smaller (1.4 to 1.6 %) with a significantly higher $\Delta AVE_{thresh} \approx 100$ ppb. Moreover, comparison of the ozonolysis experiments of IBVE with and without an excess of cyclohexane C_6H_{12} (OH scavenger) shows that the presence of an excess of C_6H_{12} reduces the SOA yield from 4.4% to below 1% in the ozonolysis of IBVE. All values obtained for the AVE in this work are in good agreement with those obtained by Klotz et al. (2004) for methyl vinyl ether MVE using the same methods (Table 1).

Saturation vapor pressures of the SOA forming organic compounds can be estimated by two different methods based on the results in Table 1. The partitioning coefficient K_i ($\mu g/m^3$) of the hypothetical semi-volatile compound i is inversely proportional to its saturation vapor pressure P_i° (Torr) according to the following relationship (Eq. 5) (Odum et al., 1996; Pankow, 1994):

$$P_i^\circ = (760 RT) / (MW_i 10^6 K_i \zeta_i) \quad (5)$$

with $R = 8.206 \times 10^{-5}$ ($m^3 \text{ atm} / (\text{mol K})$) and $T = 296 \text{ K}$. In the later sections, the newly formed aerosol is shown by ESI(+)/MS to consist mainly of oligomers in the mass

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range 300 to 500 u. Assuming that the hypothetical semi-volatile compound i represents an average of these oligomers, and that other unidentified SOA constituents are minor, a mean molar weight $MW_i = 400 \text{ g/mol}$ is assumed. With the elemental compositions of the oligomers pointing out to an oligoperoxidic structure (see later sections), the compound i is assumed to be moderately polar, giving an activity coefficient $\zeta_i \approx 1$. Another method to estimate the mean saturation vapor pressures P_{vap} [$\mu \text{ Torr}$] of the main SOA constituents relies on the measured yields of these constituents, here assumed to be nearly equal to the measured SOA yields a , and the SOA formation thresholds ΔAVE_{thresh} [ppb] (Eq. 6) (Koch et al., 2000):

$$P_{vap} = a \times \Delta AVE_{thresh} \times 10^{-9} \times P_{total} \times 10^6 \quad (6)$$

with $P_{total} = 730 \text{ Torr}$. Both methods give saturation vapor pressures between 0.1 and $1.8 \mu \text{ Torr}$, thus within one order of magnitude.

These values can only be considered as an estimate. Both calculations are based on liquid compounds with flat surfaces and do not take into account the effects of Kelvin's law, which are expected to be significant for the observed particle sizes of several nanometers. The real saturation vapor pressures of the SOA forming compounds might thus be much lower, the range of values given by the two above-described calculation methods representing upper limits.

3.3 Chemical analysis of the SOA composition and formation of oligomers

3.3.1 Observation of oligomers by ESI(+)/MS-TOF

Deploying the smooth ionisation of the electrospray technique, oligomeric products were detected in the SOA filter samples for all the enol ethers studied. Figures A1a–d (see Appendix A, <http://www.atmos-chem-phys-discuss.net/6/5629/2006/acpd-6-5629-2006-supplement.pdf>) show the mass spectra of the aerosol samples obtained during the ozonolysis of EVE, EPE, PVE and TBVE. The spectra show the presence of ions in the range between m/z 200 and 800 with regular differences $\Delta m/z$.

The pseudomolecular ions are positively charged of the type $[M+H]^+$, where M is the mass of the molecular species.

Detailed analysis of the ion spectra revealed that the ion peaks could be grouped in series, whose ions display regular differences of $\Delta m/z=46$ for all the investigated alkyl vinyl ethers AVE ($ROCH=CH_2$), and $\Delta m/z=60$ for the ethyl propenyl ether EPE ($C_2H_5OCH=CHCH_3$). Such a regular structure is typical for oligomers, consisting of a chain unit with mass 46 and 60, respectively.

The different ion series for the various unsaturated ethers are presented in Table 2 and correspond to different types of oligomers designated (a), (b), (c), (d), (e) for the AVE and (a), (B), (C), (D), (E) for EPE. In order to better distinguish the series, the peaks of the series are identified by different colours, corresponding to the colours of the peaks of Fig. A1a–d. The most intense series observed for each ether is coloured in red, and is assigned as oligomer of type (a). As described in a later section of this publication, ESI(+)/MS/MS-TOF experiments allowed to fragment the pseudomolecular ions $[M+H]^+$ and thus to determine the minimum number n of chain units 46 or 60 contained in the molecular species. In Figs. A1a–d, n are given for the pseudomolecular ions $[M+H]^+$ of the most intensive oligomer series designated as type (a). Ions of type (a) and ions of weaker oligomer series suggested to carry similar numbers of chain units, are arranged in vertical columns in Table 2.

Some interesting relationships between the oligomer ions shown in Table 2 and Figs. A1a–d and the original structure of the enol ethers can be derived. The schematic structure of an enol ether is shown in Fig. 6 and displays the two alkyl substituents, R' on the vinylic side $=CHR'$ of the double bond and R on its enolic side $=CHOR$

The substituents R and R' vary among the different ethers studied in this work. R'=H for the five alkyl vinyl ethers (AVE, $ROCH=CH_2$), while R'=CH₃ for ethyl propenyl ether (EPE, $C_2H_5OCH=CHCH_3$). R' differs thus by mass 14 (CH₂) between the AVE and EPE, a difference that is reflected in the masses of the observed oligomer chain units, which are of $\Delta m/z=46$ for the AVE and of $\Delta m/z=60=46+14$ for EPE. This indicates that the oligomer chain unit might depend on the alkyl substituents on the vinylic side

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$=CHR'$ of the double bond.

A comparison of the oligomers of type (a) formed from the five AVE ($ROCH=CH_2$) shows that their masses differ by as many multiples of mass 14 (CH₂) as do the original alkyl vinyl ethers among each other. As can be seen in Table 2 the oligomers formed from the three isomers of butyl vinyl ether, BVE, IBVE and TBVE, show the same ions of type (a). Moreover, the type (a) ions are shifted by $\Delta m/z=14$ from EVE to PVE as well as from PVE to the butyl vinyl ethers. Regarding the structure of the original ethers, these shifts might be attributed to the alkyl substituents R on the enolic side $=CHOR$ of the double bond, which contain an additional CH₂ group between EVE ($C_2H_5OCH=CH_2$, R=C₂H₅) and PVE ($C_3H_7OCH=CH_2$, R=C₃H₇), as well as between PVE and the BVEs ($C_4H_9OCH=CH_2$, R=C₄H₉). Consequently, the alkoxy group OR of the original ether is suggested to be contained once in each oligomer ion.

Similar observations are made for the other, less intense oligomer series shown in Table 2 and in Figs. A1a–d. Like the oligomer ions of type (a), all weaker series also show chain units of mass 46 for AVE and of mass 60 for EPE. For all AVE except of TBVE, oligomer ions of type (b) are visible, coloured in green in Table 2 and Figs. A1a–d. They differ from the series of type (a) ions by an additional $\Delta m/z=16$, possibly corresponding to an oxygen atom. Consequently, like the type (a) ions, oligomer series of type (b) are also shifted by $\Delta m/z=14$ between EVE and PVE as well as between PVE and BVE/IBVE. The three butyl isomers show a variety of additional ions series labelled (c), (d) and (e). For example, the ion series of type (c) systematically possess an additional $\Delta m/z=14$ with respect to the type (b) oligomer series, and consequently an additional $\Delta m/z=30$ with respect to the type (a) series, eventually accounting in total for a formaldehyde-like unit (CH₂O).

The mass spectra obtained from the ozonolysis of ethyl propenyl ether (EPE) also show besides the type (a) oligomer, a variety of additional oligomer series labelled type (B) through (E). The capital letters distinguish these ion series from those formed from the AVE and designated (b) through (e), as they do not seem to show any analogy.

3.3.2 Additional test experiments

Background experiments were performed to assure that the detected oligomers are definitely produced through the ozonolysis of unsaturated ethers during the typical duration of the experiment (45 min) and substantially contribute to SOA formation. No aerosol was formed when 9 ppm of EVE was introduced alone in the reactor, nor when solely ozone (8 ppm) was introduced, and filter samples of both experiments did not show any of the described oligomer ions (Fig. A2).

SOA formation was observed to be strongly reduced or suppressed upon addition of HCOOH in excess during gas-phase ozonolysis of enol ethers (Sadezky, 2005) and terpenes (Bonn et al., 2002). The addition of 20 ppm of HCOOH to the mixture of EVE (9 ppm) and O₃ (8 ppm) reduced the aerosol yield to less than 1 µg/m³, whereas in the absence of HCOOH the typical aerosol yield was in the order of 5 mg/m³. SOA particle size distributions measured after completion of the reaction, just before the beginning of the filter sampling, are shown for both experiments in Figs. 7a and b. The mass spectrum of the collected filter samples of the experiment with added HCOOH is displayed in Fig. A2b and does not show any characteristic ions of the oligomers described before. The most intense ions visible in the Figs. A2a and b, 288 and 299, were proven by MS/MS experiments not to possess an oligomeric structure. HCOOH is known to scavenge Criegee Intermediates formed during alkene ozonolysis, leading to the formation of gaseous hydroperoxy formates (e.g. Klotz et al., 2004; Neeb et al., 1998; Horie et al., 1997; Sadezky, 2005). Considering that this reaction is responsible for the observed strong reduction of SOA formation, the scavenging effect of HCOOH indicates that Criegee Intermediates play a decisive role in SOA formation. The addition of H₂O to the ozonolysis system also caused a remarkable reduction of SOA formation (Sadezky, 2005).

In the presence of 300 ppm of cyclohexane (an OH scavenger), the total aerosol yield decreases by 80%, as is also described for IBVE in the previous section (Table 1). However, mass spectra obtained from filter samples of these experiments remained un-

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changed in comparison with those from ozonolysis experiments with mixtures of ether and ozone only, showing the same oligomer ions. The SOA particle size distribution measured after completion of the reaction, before the beginning of the filter sampling, is shown in Fig. 7c for an ozonolysis experiment of EVE (9 ppm of EVE and 8 ppm of ozone) with added 300 ppm of cyclohexane producing about 0.8 mg/m³ of SOA. For comparison, also the SOA particle size distribution measured for the reaction of a mixture of EPE (9 ppm) with ozone (8 ppm), leading to the formation of about 3 mg/m³ of SOA, is shown in Fig. 7d.

3.3.3 Separation of the oligomers by HPLC

The different oligomeric constituents of the SOA may be separated by HPLC according to their number of chain units. An example of the TIC (total ion current) chromatogram is shown in Fig. A3a, representing the SOA formed during ozonolysis of EVE, where chromatographic peaks in different retention time (RT) zones surrounded with different colours are marked. In Fig. A3b, the mass spectra of each of these four retention time (RT) zones, displayed in the corresponding colours, are overlaid. Figure A3b shows that the most intense ion present in each retention time (RT) zone represents a pseudomolecular ion of the oligomer of type (a) (see Table 2 and Fig. A1a), with $m/z=329$ for the first (blue-marked) retention time zone at RT=16 min, $m/z=375$ for the second (red-marked) at RT=17 min, $m/z=421$ for the third (green-marked) at RT=17.9 min, and $m/z=467$ for the fourth (violet-marked) at RT=18.5 min. The less intense pseudomolecular ions of the oligomer of type (b) with $m/z=345$ (blue-marked zone, RT=16 min), 391 (red-marked zone, RT=17 min), 437 (green-marked zone, RT=17.9 min) and 483 (violet-marked zone, RT=18.5 min) are also shown (see Table 2 and Fig. A1a). Thus, with increasing retention time, from one chromatographic peak to the next, the masses of the detected oligomeric pseudomolecular ions increase by $\Delta m/z=46$, that is, by the mass of a chain unit.

Similar TIC chromatograms and corresponding mass spectra were obtained for the SOA formed from ozonolysis of the other AVEs, where the oligomers all differ with the

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chain unit $\Delta m/z=46$ between two consecutive chromatographic peaks, and for EPE, where the chain unit is $\Delta m/z=60$. For all enol ethers, retention times of the different oligomers strictly increase with increasing number of chain units. A straight increase or decrease of the retention time with the number of chain units is a characteristic behaviour of polymeric compounds with oxygenated chain carriers, such as ethylene oxide, in reverse-phase liquid chromatography (Chaimbault et al., 1998).

The peaks at retention times (RT) near 16 and 17 min (blue-marked and red-marked zones) of the TIC chromatogram of Fig. A3a can be resolved in the XIC (eXtracted Ion Current) mode as shown in Fig. A3c. The blue traces correspond to $m/z=329$ and 375 , respectively of the oligomer of type (a), whereas the red traces correspond to $m/z=345$ and 391 , of the oligomer of type (b). Four peaks are visible for each ion, which indicates that at least four isomers of the types (a) and (b) exist. Similar observations were made for the ions with $m/z=343$ and 389 (type a) and $m/z=359$ and 405 (type b) for PVE, and for the ions with $m/z=375$ and 403 (type a) and $m/z=373$ and 419 (type b) for *n*-BVE.

3.3.4 Fragmentation of oligomers using ESI(+)/MS-MS-TOF

The fragmentation of the pseudomolecular ions using the MS/MS mode confirm that they consist of an oligomeric structure with the chain unit of $\Delta m/z=60$ for SOA formed in the ozonolysis of ethyl propenyl ether (EPE) and $\Delta m/z=46$ for SOA formed from all the alkyl vinyl ethers (AVE) studied. The majority of the observed fragment ions can be classified as being part of fragmentation pathways, formed by successive loss of the chain unit as neutral mass 60 or 46 for EPE and AVE, respectively. Tables 3a and b summarize the observed ions of the fragmentation pathways for the oligomers of type (a), respectively involving loss of chain units $\Delta m/z=46$ for AVE (Table 3a), and loss of chain units $\Delta m/z=60$ for EPE (Table 3b). A variety of nine fragmentation pathways can thus be distinguished for an oligomer of type (a).

Generally, for the different AVE, detailed analysis of the MS/MS spectra has lead to the observation that the fragmentation of their oligomers proceeds according to a

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common scheme. There exist two distinct fragmentation patterns involving the loss of the chain element $\Delta m/z=46$. One produces series of fragmentations in a way that the RO group clearly remains in the fragment ions. These fragment ions, given in the central column of Table 3a, differ by $\Delta m/z=14$ between EVE and PVE, as well as between PVE and the three BVEs, as do the pseudomolecular ions themselves. The most intense fragmentation pathway showing this pattern is marked in violet colour, and consists a.o. of the ions with m/z 127, 173, 219, 265, etc. for EVE (molar weight MW: 72 g/mol), the ions with m/z 141, 187, 233, 279, etc. for PVE (MW: 86 g/mol), and the ions with m/z 155, 201, 247, 293, etc. for the three BVE-isomers (MW: 100 g/mol). The other fragmentation pattern produces fragment ions, which do not contain the RO group and thus are identical for all the AVE. Those fragment ions are listed in the right-hand column of Table 3a. The two most important pathways showing this pattern are represented by red and green colours, and consist a.o. of the ions with m/z 85, 131, 177, 223, etc. and 99, 145, 191, 237, etc., respectively.

Examples of the MS/MS spectra are displayed in Figs. A4a–d for various unsaturated ethers, which show the fragmentation of a selected specific pseudo-molecular ion of the type (a) oligomer. The three fragmentation pathways highlighted in violet, red and green colours in Table 3a and b are shown in the MS/MS spectra in Figs. A4a–d and are also schematically displayed in the bottom section of these figures by use of the same colours. The elemental compositions given for the ions and neutral molecules are discussed below. Note that for AVE and EPE, the fragmentation pathways marked in red and green terminate by fragment ions of similar m/z , 85 and 99, respectively, and that for EVE and EPE, the fragmentation pathways marked in violet both terminate with fragment ions of a similar m/z , 127.

3.3.5 Principal structure of the oligomer (a)

The fragmentation pathways characterized by the regular loss of the chain unit with $\Delta m/z=60$ for EPE and $\Delta m/z=46$ for the AVE typically range from $m/z=300$ to 400 , the masses of the pseudo-molecular ions, down to m/z about 100 for the smallest

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fragment ions (Tables 3a and b). Thus, they reveal a linear oligomeric structure, which can schematically be represented for an oligomer of type (a) by a chain having a starting group "X" and an end group "Y" linked by several chain units of mass 60 or 46 for EPE and AVE, respectively. This structure is schematically pictured in Fig. 8 for an oligomer of type (a) formed from AVE, with chain units of mass 46. The fragmentation initiates with a loss of a neutral species X or Y from one side of the pseudo molecular ion $[X-(46)_n-Y+H]^+$, and continues with n successive losses of a chain unit of mass 46, to finally terminate on a fragment ion XH^+ or YH^+ , respectively, representing the opposite side of the molecule.

Attribution of the alkoxy group RO to either the starting group X or the end group Y of the pseudo molecular ion $[X-(46)_n-Y+H]^+$ explains the two distinct fragmentation patterns in the MS/MS spectra of the type (a) oligomers formed from AVE. In Fig. 8, RO is for example proposed to be part of the starting group X. The first fragmentation pattern, which conserves the RO group in the fragment ions, terminates with the fragment ion XH^+ , as shown on the right-hand side of Fig. 8. Application of this schema to the observed fragmentation pathway marked in violet colour (Table 3a and Figs. A4a, c and d) identifies the terminating fragment ions 127 (EVE), 141 (PVE) and 155 (the three isomers of BVE) as XH^+ and thus as the starting group X of the pseudo molecular ion $[X-(46)_n-Y+H]^+$. Consequently, in this example, the end group Y is represented by a neutral species of mass 64, whose loss from the pseudo molecular ion $[X-(46)_n-Y+H]^+$ initiates the violet-coloured fragmentation pathway for type (a) oligomers from all AVE (Figs. A4a, c and d). The second fragmentation pattern, whose fragment ions do not contain the RO group and are identical for type (a) oligomers formed from different AVE, terminates on the fragment ion YH^+ , as shown on the left-hand side of Fig. 8. The starting group X, which contains the RO group, is fragmented as neutral species from the pseudo molecular ion $[X-(46)_n-Y+H]^+$ in the first step of this fragmentation pattern. Application of this scheme to the observed fragmentation pathways marked in red and green colours attributes their terminating fragment ions 85 and 99 to YH^+ , thus to the end group Y of the pseudo molecular ion $[X-(46)_n-Y+H]^+$. The starting group X

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is consequently represented by the neutral species first fragmented from the pseudo-molecular ion $[X-(46)_n-Y+H]^+$, which are of masses 92 and 106 (EVE), 106 and 120 (PVE), and 120 and 134 (the three isomers of BVE) for the green- and red-coloured fragmentation pathways, respectively (Figs. A4a, c and d), and thus show the expected differences in units of $\Delta m/z=14$ according to the contained RO groups.

Separation of the oligomers by liquid chromatography proved the existence of at least four isomers of the pseudomolecular ion, as is shown in Fig. A3c for oligomers formed during the ozonolysis of EVE. The presence of these isomers explains the complexity of the MS/MS spectra, with at least nine identified fragmentation pathways (Table 3a) belonging to one of the two fragmentation patterns. The regularity of the different fragmentation pathways and the strict repetition of the two described fragmentation patterns suggest that the isomers follow the same structural principle. Therefore, in Fig. 8, isomers of the pseudomolecular ion $[X-(46)_n-Y+H]^+$ are symbolized by the addition of two more pseudomolecular structures $[A-(46)_n-Z+H]^+$ and $[B-(46)_n-W+H]^+$. As is schematically shown in Fig. 8, MS/MS spectra thus contain the fragmentation pathways of all isomers $[X-(46)_n-Y+H]^+$, $[A-(46)_n-Z+H]^+$ and $[B-(46)_n-W+H]^+$ of a pseudo-molecular ion according to the two fragmentation patterns depicted in Fig. 8. Note that the proton ionizing the oligomer molecule is suggested in Fig. 8 to remain attached to the fragment ions. This is however an arbitrary decision for illustration purposes, as the determination of the exact fragmentation mechanism needs further investigation.

3.3.6 Elemental composition and minimum number n of the chain units 46 and 60

Possible elemental compositions of parent ions, fragment ions and fragmented neutral molecules were determined from the m/z values measured during MS/MS experiments. Resulting elemental compositions are given exemplarily in Figs. A4a–d for those ions and neutrals that are part of one of the three fragmentation pathways demonstrated in violet, red and green colours in the MS/MS spectra of oligomer (a) ions. Detailed results of the calculations are given for two of these pathways in Tables 4a and b: for the red-coloured fragmentation of the pseudomolecular ion 399

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formed from EPE (see MS/MS spectrum in Fig. A4b) and for the violet-coloured fragmentation of the pseudomolecular ion 389 formed from PVE (MS/MS spectrum in Fig. A4c). Similar calculations from MS/MS spectra of several other pseudomolecular oligomeric ions formed from the diverse AVE and from EPE, for the three green-
5 red- and violet-coloured fragmentation pathways, as well as for other fragmentation pathways listed in Tables 3a and b, are given in Sadezky (2005) and show the reproducibility of these results.

Measured m/z values allow a rather doubtless determination of the elemental compositions of the fragmented neutral molecules, in particular of the chain units 46 and 60.
10 Analysis of the calculated mass errors in Tables 4a and b show a strong and systematic preference of the elemental compositions CH_2O_2 (theoretical mass 46,0046 u) for the chain unit of mass 46 (AVE) and $\text{C}_2\text{H}_4\text{O}_2$ (theoretical mass 60,0211 u) for the chain unit of mass 60 (EPE). The only other elemental compositions, which make chemical sense as neutral molecules in order to represent the chain units, $\text{C}_2\text{H}_6\text{O}$ for mass 46
15 and $\text{C}_3\text{H}_8\text{O}$ for mass 60, usually show much larger mass errors.

Further work is needed to elucidate the structures of parent and fragment ions. For their chemical compositions, the elements C, H and O and a single positive charge are taken into account. The fact that good fragmentation spectra are obtained from pseudomolecular ions during MS/MS experiments disagree with Na^+ adducts.

20 The existence of different isomers of the pseudomolecular ions further complicate a precise determination of the oligomer structures. In the MS/MS spectra, according to Fig. 8, fragment ions representing the respective starting and end groups A, B, X, Z, W, and Y of all isomers of the pseudomolecular ion appear together. Even for the alkyl vinyl ethers AVE, for which the fragment ions containing the starting groups X, A, and B and those containing the end groups Y, Z and W can be distinguished, fragment ions
25 of different pathways cannot clearly be attributed to specific isomers so far. It is thus, for example, unknown whether two of the three fragmentation pathways discussed in detail in the sections above and demonstrated in violet, red and green colours (Table 3a and b and Figs. A4a–d) are complementary and generated by the same isomer, or if

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all three pathways originate from different isomers.

MS/MS experiments allow to determine the minimum number of chain units n comprised in an oligomer molecule. n is derived from the number of successive losses of the chain unit as a neutral molecule which are directly observed in the MS/MS spectra.
5 Its values are given in Table 2 and in the MS spectra of the oligomers in Figs. A1a–d, and are between 2 and 9 for most oligomers. The most abundant pseudomolecular ions contain 3 to 4 chain units.

4 Discussion

It might be essential to consider the role of the Criegee Intermediate (CI) in the formation of oligomers. Indeed, the linking chain units observed in this study of the different
10 oligomers have the same elementary composition as the CIs: CH_2O_2 ($=\text{CH}_2\text{OO}$ for $\text{C}_1\text{-CI}$) of mass 46 for the alkyl vinyl ethers (AVE) and $\text{C}_2\text{H}_4\text{O}_2$ ($=\text{CH}_3\text{CHOO}$ for $\text{C}_2\text{-CI}$) of mass 60 for ethyl propenyl ether (EPE).

The structure of the oligomer, which is schematically shown in Fig. 8, is deduced from the fragmentation pattern of the main ions, obtained by ESI(+)/Q-TOF MS/MS, as is displayed in Figs. A4a–d. The fragments, whose successive and systematic loss is observed in the fragmentation spectra, are neutral species with the chemical compositions CH_2O_2 (mass 46) and $\text{C}_2\text{H}_4\text{O}_2$ (mass 60), respectively for the AVE and EPE.
15 Based on these observations, we suggest that the oligomer might have the following basic structure of an oligoperoxide shown in Fig. 9:

Moreover, the ozonolysis experiments performed in the presence of HCOOH , which is known to be an efficient Criegee Intermediate scavenger, have shown that the yield of SOA formation was drastically reduced, eventually suppressed, and that oligomers in detectable quantities are not formed. SOA formation is also strongly reduced in the
25 presence of H_2O (Sadezky, 2005). Consequently, this reaction obviously competes with known reactions of the stabilized CI (SCI) with HCOOH or H_2O , which lead to the formation of small gaseous peroxide compounds (Calvert et al., 2000). Oligoperoxide

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formation is thus another, so far unknown reaction of stabilized SCIs in the gas phase leading to SOA formation.

Oligomers have long time been recognized to be significant products of the ozonolysis in solutions (Barton et al., 2004; Lockley et al., 2001; and references therein).
5 Early studies by Criegee (1957) speculated on the products of the ozonolysis incorporated in the oligomeric structures, namely the CI, also called carbonyl oxide. Lockley et al. (2001) investigated the ozonolysis of tetramethylethylene (TME) at -60°C and identified acetone cyclic di- and triperoxides, and substantial amounts of oligomeric peroxides of acetone carbonyl oxide. Both linear and cyclic oligomeric peroxides were
10 inferred from GC-MS analysis. Also recently Barton et al. (2004) performed the ozonolysis of tetramethylethylene (TME) in pentane as a nonparticipating solvent at -60°C yielding predominantly oligoperoxides with m/z up to 900 u, as identified by ESI/MS under soft ionization conditions. The peroxidic chain link of the mass peaks was of mass 74, which corresponds to the mass of the acetone carbonyl oxide $(\text{CH}_3)_2\text{COO}$,
15 thus the CI formed in the ozonolysis of TME (Fig. 10). Up to 12 chain units were observed for different oligoperoxides. Moreover, these oligomeric compounds were found to be thermally stable when extracted and refluxed in methanol for a few hours (Barton et al., 2004).

In liquid phase, the formation of small peroxidic compounds in the presence of protonic solvents or additives, such as H_2O , alcohols and organic acids, is also well known
20 (e.g. Bunnelle, 1991; Thamm et al., 1996). In the liquid-phase ozonolysis of TME, oligoperoxide formation was only observed in non-protonated solvents (pentane) by Barton et al. (2004). The suppression of SOA and oligomer formation in the presence of HCOOH observed in the gas phase in this work is in analogy to the liquid-phase
25 chemistry.

It is to be concluded that the CIs play a pertinent role in the oligomerization also in the gas phase. The mechanism of formation of the oligoperoxides found in the ozonolysis of enol ethers remains an enigma, since this type of repetitive chain structure involving CH_2OO and CH_3CHOO units has not been observed previously in gas phase reactions.

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Several studies have reported the formation of SOA in the gas phase oxidation of VOC and identified polymerization reactions in the presence (Tolocka et al., 2004; Gao et al., 2004) and absence (Kalberer et al., 2004; Ziemann, 2002, 2003; Zahardis et al., 2005) of preexisting acid seed aerosols. Most of these studies invoked oligomerization
5 by acid-catalysed aldol condensation and *gem*-diol formation of main carbonyl products of the ozonolysis reaction. Tolocka et al. (2004) reported oligomers in the SOA produced in the ozonolysis of α -pinene, which are formed from such reactions of pinonaldehyde. Gao et al. (2004) observed formation of large oligomers (up to m/z 1600) in the ozonolysis of a series of cyclic alkenes, and reported similar condensation
10 reactions among carbonyl reaction products. Kalberer et al. (2004) detected oligomers in SOA formed during photooxidation of 1,3,5-trimethylbenzene, and attributed them to hydration-condensation reactions involving the main reaction products glyoxal and methylglyoxal. Peroxidic products were taken into consideration by other previous studies: Ziemann (2002) suggested the formation of diacylperoxides that act as nucleating
15 agents in the ozonolysis of cyclic alkenes. A mechanism was presented involving the cross reaction of acylperoxy radicals, formed upon isomerisation and decomposition of excited Criegee Intermediates in air. In the presence of alcohols Ziemann (2003) invoked the formation of alkoxyhydroperoxy aldehydes, which were generated from the reaction of SCI with alcohols, and subsequently undergo intermolecular reaction
20 to form cyclic peroxyhemiacetals. The only other study that so far reported reactions among Criegee Intermediates was recently published by Zahardis et al. (2005). The authors studied the ozonolysis of oleic acid and identified secondary ozonides and cyclic geminal diperoxides, the latter being formed through the recombination of two SCI.

25 5 Conclusion

In this work, we investigated secondary organic aerosol (SOA) formation from the gas phase ozonolysis of enol ethers. In spite of the fact that these compounds consist of

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only 5 to 7 carbon atoms, significant SOA formation was observed. Chemical analysis of the formed SOA allowed us to suggest a new reaction mechanism for new SOA particle formation: the oligomerization of Criegee Intermediates leading to the formation of oligoperoxidic chains with molecular weights up to 300 to 500 u. The suggested mechanism is in agreement with the formation of oligoperoxidic compounds from Criegee Intermediates, which was observed during liquid phase ozonolysis in earlier studies. Thus, in this work, we report the first observation of a similar reaction during gas phase ozonolysis. It is suggested that in the gas phase, oligoperoxides are formed from reactions of stabilized Criegee Intermediates, the detailed mechanisms of these reactions, however, are so far unknown and need further investigation.

Earlier studies have shown that the main tropospheric degradation mechanisms of unsaturated ethers are their reactions with ozone and OH radicals (Thiault et al., 2002). The present study shows that mixing ratios of a few hundred ppb of enol ether and ozone are sufficient for significant formation of SOA. However, stabilized Criegee Intermediates are expected to react mainly with water vapour in the troposphere to form gaseous hydroxyhydroperoxides (Calvert et al., 2000), a reaction that has been shown to compete with other known reactions of stabilized Criegee Intermediates as well as with SOA formation (Sadezky, 2005; Bonn et al., 2002). The relevance of the observed oligoperoxide formation for SOA formation in the atmosphere has thus to be further investigated. In contrast to the other mechanisms of oligomer formation suggested to contribute to tropospheric SOA formation, which always involve heterogeneous or multiphase reactions on preexisting aerosol particles (e.g. Tolocka et al., 2004; Gao et al., 2004), the oligoperoxide formation observed in the present work takes place in the gas phase and obviously leads to new SOA particle formation by homogeneous nucleation of its reaction products.

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Table 1. SOA formation during ozonolysis of alkyl vinyl ethers (AVE): SOA mass yields a (Eq. 1) and SOA formation thresholds ΔAVE_{thresh} in comparison with the yields α_i and partitioning coefficients K_i of a hypothetical semi-volatile SOA constituent i (Eq. 4), and estimated saturation vapour pressures P_i° (Eq. 5) and P_{vap} (Eq. 6) of main SOA constituents.

AVE	a (%)	SOA formation threshold ΔAVE_{thresh}	α_i (%)	K_i (m ³ /μg)	P_i° (μTorr)	P_{vap} (μTorr)
EVE*	3.5±0.2	(173±6) μg/m ³ (58±2) ppb	3.5±0.1	0.17±0.02	0.27	1.48
PVE*	3.8±0.2	(190±6) μg/m ³ (54±2) ppb	3.8±0.2	0.14±0.02	0.33	1.50
BVE*	4.3±0.4	(223±45) μg/m ³ (54±11) ppb	4.4±0.2	0.10±0.01	0.46	1.70
TBVE*	1.6±0.3	(407±88) μg/m ³ (100±22) ppb	1.4±0.2	0.20±0.06	0.23	1.17
IBVE*	4.4±0.1	(151±19) μg/m ³ (37±5) ppb	4.7±0.2	0.13±0.01	0.36	1.19
IBVE	0.9±0.1	(202±7) μg/m ³ (49±1) ppb	0.9±0.1	0.60±0.08	0.08	0.32
+ excess of C ₆ H ₁₂ * MVE [#]	0.66±0.07	(162±38) μg/m ³ 68 ppb	0.66±0.13	0.98±0.27	0.05	0.33

* this work
Klotz et al. (2004)

Table 2. Oligomer pseudomolecular ion series detected by ESI(+)/MS-TOF in the SOA formed during ozonolysis of enol ethers (n: minimum number of chain units determined by MS/MS experiments) (MW: molar weight [g/mol]).

Ether	Ion Series (m/z)										
n	n=1	n=2	n=3	n=4	n=5	n=6	n=7	n=8	n=9	n = 10	
Ethyl Propenyl Ether (MW 86 g/mol)		339	399	459	519	579	639	699	759	(a)	
			357	417	477					(B)	
EPE, C ₂ H ₅ OCH=CHCH ₃	265	325	385	445						(C)	
	293	353	413	473	533					(D)	
				489	549	609	669	729	789	(E)	
Ethyl Vinyl Ether (MW 72g/mol)		283	329	375	421	467	513			(a)	
			345	391	437	483				(b)	
EVE, C ₂ H ₅ OCH=CH ₂											
Propyl Vinyl Ether (MW 86 g/mol)		297	343	389	435	481	527	573		(a)	
			359	405	451	497				(b)	
PVE, C ₃ H ₇ OCH=CH ₂											
n-Butyl Vinyl Ether (MW 100 g/mol)			357	403	449	495	541	587	633	679	(a)
					465	511	557	603	649	695	(b)
BVE, C ₄ H ₇ OCH=CH ₂	295	341	387	433	479	525	571	617	663	(c)	
	293	339	385	431	477	523				(d)	
iso-Butyl Vinyl Ether (MW 100 g/mol)		311	357	403	449	495	541	587	633	(a)	
			373	419	465	511	567	603		(b)	
IBVE, (CH ₃) ₂ CHCH ₂ OCH=CH ₂		341	387	433	479	525	571			(c)	
	247	293	339	385	431					(d)	
t-Butyl Vinyl Ether (MW 100 g/mol)	265	311	357	403	449	495	541			(a)	
	295	341	387	433	479					(c)	
TBVE, (CH ₃) ₃ COCH=CH ₂	247	293	339	385						(d)	
					427	473	519	565	611	(e)	

Table 3. (a) Ions of fragmentation pathways consisting of successive loss of chain unit $\Delta m/z46$, with or without conservation of the ether group RO, which were observed for the oligomer of type (a) formed in the ozonolysis of alkyl vinyl ethers (AVE) (MW: molar weight [g/mol]).

Ether Parent Ions (m/z)	Fragment Ion series (m/z) conserving RO	Fragment Ion series (m/z) not-conserving RO
Ethyl Vinyl Ether (MW 72 g/mol) (EVE, $C_2H_5OCH=CH_2$) 329, 375, 421, 467	403, 357, 311, 265, 219, 173, 127 339, 293, 247, 201, 155 374, 328, 282, 236, 190	
Propyl Vinyl Ether (MW 86 g/mol) (PVE, $C_3H_7OCH=CH_2$) 343, 389, 435, 481	417, 371, 325, 279, 233, 187, 141 353, 307, 261, 215, 169 388, 342, 296, 250, 204	253, 207, 161, 115 313, 267, 221, 175, 129, 83 315, 269, 223, 177, 131, 85
<i>n</i> -Butyl Vinyl Ether (MW 100 g/mol) (BVE, $C_4H_9OCH=CH_2$) 357, 403, 449, 495	} 431, 385, 339, 293, 247, 201, 155 367, 321, 275, 229, 183 402, 356, 310, 264, 218	225, 179, 133 329, 283, 237, 191, 145, 99 286, 240, 194, 148, 102
<i>iso</i> -Butyl Vinyl Ether (MW 100 g/mol) (IBVE, $(CH_3)_2CHCH_2OCH=CH_2$) 357, 403, 449, 495		
<i>t</i> -Butyl Vinyl Ether (MW 100 g/mol) (TBVE, $(CH_3)_3COCH=CH_2$) 357, 403, 449, 495		

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Table 3. (b) Ions of fragmentation pathways consisting of successive loss of chain unit $\Delta m/z60$, which were observed for the oligomer of type (a) formed in the ozonolysis of ethyl propenyl ether EPE (MW: molar weight [g/mol]).

Ether Parent Ions (m/z)	Fragment Ion series (m/z)
	367, 307, 247, 187, 127 325, 265, 205, 145, 85 339, 279, 219, 159, 99
Ethyl Propenyl Ether (MW 86 g/mol) (EPE, $C_2H_5OCH=CHCH_3$) 279, 339, 399, 459	381, 321, 261, 201, 141 293, 233, 173 341, 281, 221, 161 249, 189, 129 297, 237, 177, 117 296, 236, 176

5658

Table 4. (a) Calculated elemental compositions of parent ions, fragment ions and fragmented neutral molecules for the red-coloured fragmentation pathway of the pseudomolecular oligomer ion 399 formed from EPE (see MS/MS spectrum in Fig. A4b).

Measured mass (u)	Elemental composition	Exact mass (u)	Absolute mass error (mDa)	Relative mass error (ppm)
399,1556 parent ion [X-[60] ₂ -Y + H] ⁺	C ₁₈ H ₂₇ O ₁₂ C ₁₈ H ₃₁ O ₁₁ C ₁₄ H ₂₃ O ₁₃ C ₁₇ H ₃₅ O ₁₀	399,1497 399,18608 399,11331 399,22247	5,8968 -30,4886 42,2823 -66,8742	14,7732 -76,3828 105,9293 -167,5389
134,1036 initially fragmented neutral X or Y	C ₁₀ H ₁₄ C ₇ H ₄ O ₂ C ₉ H ₁₀ O C ₉ H ₁₀ O ₄	134,10955 134,09429 134,07316 134,0579	-5,9505 9,3055 30,4349 45,6910	-44,3728 69,3905 226,9511 340,7144
265,0520 fragment ion [Y-[60] ₂ + H] ⁺ or [X-[60] ₂ + H] ⁺	C ₈ H ₁₃ O ₉ C ₈ H ₉ O ₁₀ C ₁₀ H ₁₇ O ₈ C ₁₁ H ₂₁ O ₇	265,0554 265,01902 265,09179 265,12817	-3,4086 32,9768 -39,7942 -76,1797	-12,8604 124,4161 -150,1370 -287,4137
60,0149 neutral chain carrier 60	C ₂ H ₄ O ₂ CO ₃ C ₃ H ₆ O	60,02112 59,98474 60,05751	-6,2294 30,1560 -42,6149	-103,7982 502,4765 -710,0729
205,0326 fragment ion [Y-[60] ₂ + H] ⁺ or [X-[60] ₂ + H] ⁺	C ₇ H ₆ O ₇ C ₈ H ₆ O ₈ C ₈ H ₁₃ O ₆ C ₉ H ₁₇ O ₅	205,03427 204,99789 205,07066 205,10705	-1,6792 34,7062 -38,0647 -74,4503	-8,1901 169,2714 -185,6518 -363,1134
60,0202 neutral chain carrier 60	C ₂ H ₄ O ₂ CO ₃ C ₃ H ₆ O	60,02112 59,98474 60,05751	-0,9294 35,4560 -37,3149	-15,4854 590,7357 -621,7066
145,0178 fragment ion [Y-[60] + H] ⁺ or [X-[60] + H] ⁺	C ₈ H ₆ O ₅ C ₈ H ₆ O ₄ C ₈ HO ₆ C ₇ H ₁₃ O ₃	145,0131 145,0495 144,9768 145,0859	4,6501 -31,7353 41,0356 -68,1208	32,0661 -218,8367 282,9690 -469,7395
60,0247 neutral chain carrier 60	C ₂ H ₄ O ₂ C ₃ H ₆ O CO ₃	60,02112 60,05751 59,98474	3,5705 -32,8149 39,9560	59,4848 -546,6909 665,6606
84,9931 fragment ion [Y + H] ⁺ or [X + H] ⁺	C ₃ HO ₃ C ₄ H ₆ O ₂ C ₅ H ₆ O	84,9920 85,0284 85,06479	1,0796 -35,3058 -71,6914	12,7023 -415,3945 -843,4914

5659

Table 4. (b) Calculated elemental compositions of parent ions, fragment ions and fragmented neutral molecules for the violet-coloured fragmentation pathway of the pseudomolecular oligomer ion 389 formed from PVE (see MS/MS spectrum in Fig. A4c).

Measured mass (u)	Elemental composition	Exact mass (u)	Absolute mass error (mDa)	Relative mass error (ppm)
389,0984 parent ion [X-[46] ₂ -Y + H] ⁺	C ₁₂ H ₂₁ O ₁₄ C ₁₃ H ₂₅ O ₁₃ C ₁₁ H ₁₇ O ₁₅ C ₁₄ H ₂₉ O ₁₂	389,09258 389,12896 389,05619 389,16535	5,8177 -30,5677 42,2033 -66,9532	14,9519 -78,5603 108,4641 -172,0725
64,0187 initially fragmented neutral Y	CH ₂ O ₃ C ₃ H ₄ C ₄ O O ₄	64,01604 64,0313 63,99491 63,97965	2,6559 -12,6001 23,7853 39,0414	41,4866 -196,8199 371,5376 609,8443
325,0797 fragment ion [X-[46] ₂ + H] ⁺	C ₁₁ H ₁₇ O ₁₁ C ₁₂ H ₂₁ O ₁₀ C ₁₀ H ₁₃ O ₁₂	325,07653 325,11292 325,04015	3,1618 -33,2236 39,5473	9,7263 -102,2014 121,6542
46,0064 neutral chain carrier 46	CH ₂ O ₂ C ₂ H ₆ O	46,00547 46,04186	0,9206 -35,4648	20,0111 -770,8684
279,0733 fragment ion [X-[46] ₂ + H] ⁺	C ₁₀ H ₁₅ O ₉ C ₁₁ H ₁₉ O ₈ C ₉ H ₁₁ O ₁₀	279,07105 279,10744 279,03467	2,2412 -34,1443 38,6267	8,0309 -122,3486 138,4104
46,0093 neutral chain carrier 46	CH ₂ O ₂ C ₂ H ₆ O	46,00547 46,04186	3,8206 -32,5648	83,0406 -707,7890
233,0640 fragment ion [X-[46] ₂ + H] ⁺	C ₈ H ₁₅ O ₇ C ₈ H ₉ O ₈ C ₁₀ H ₁₇ O ₆	233,06557 233,02919 233,10196	-1,5794 34,806 -37,9649	-6,7767 149,341 -162,8945
45,9996 neutral chain carrier 46	CH ₂ O ₂ C ₂ H ₆ O	46,00547 46,04186	-5,8793 -42,2648	-127,8132 -918,8097
187,0644 fragment ion [X-[46] + H] ⁺	C ₈ H ₁₁ O ₅ C ₉ H ₁₅ O ₄ C ₇ H ₇ O ₆	187,0601 187,09648 187,02371	4,2999 -32,0855 40,6854	22,9863 -171,521 217,4937
46,0120 neutral chain carrier 46	CH ₂ O ₂ C ₂ H ₆ O	46,00547 46,04186	6,5206 -29,8648	141,716 -649,0671
141,0524 fragment ion [X + H] ⁺	C ₇ H ₉ O ₃ C ₆ H ₅ O ₄ C ₈ H ₁₃ O ₂	141,05462 141,01823 141,0910	-2,2207 34,1648 -38,6062	-15,7437 242,2127 -273,7001

5660

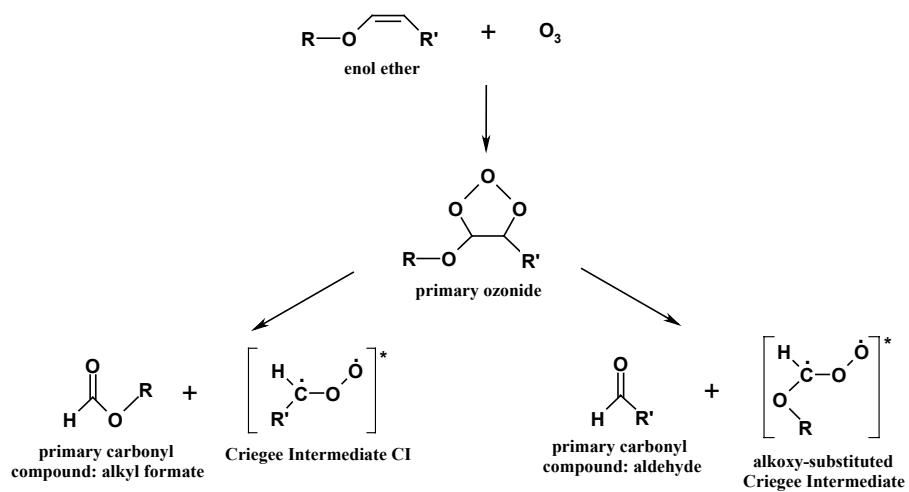


Fig. 1. General mechanism of the ozonolysis of enol ethers.

5661

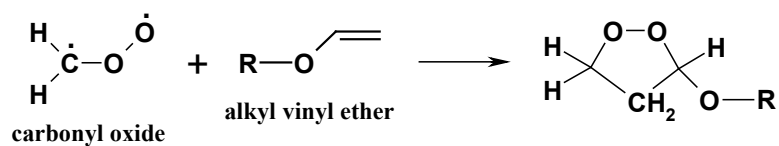


Fig. 2. 1,3-Cycloaddition of the C₁-CI into the double bond of a vinyl ether.

5662

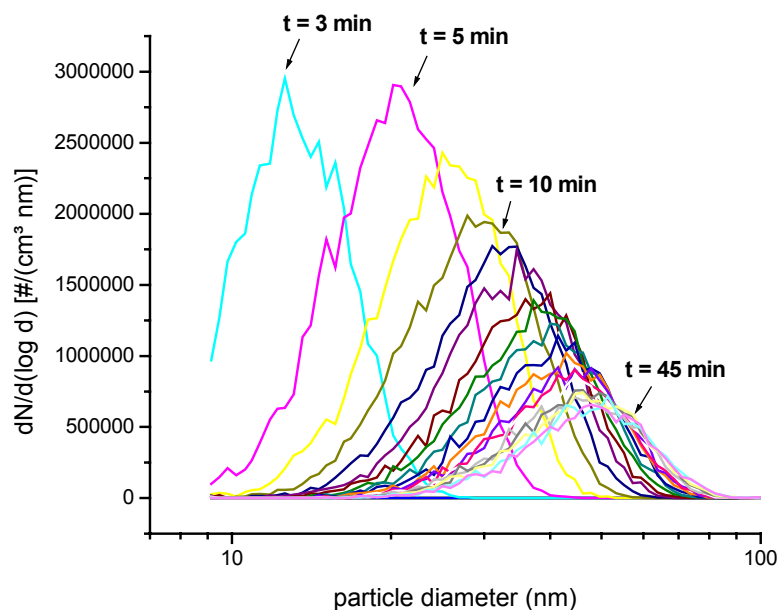


Fig. 3. Evolution of the SOA size distribution with reaction time (t) during ozonolysis of EVE (initial reactant mixing ratios: 300 ppb).

5663

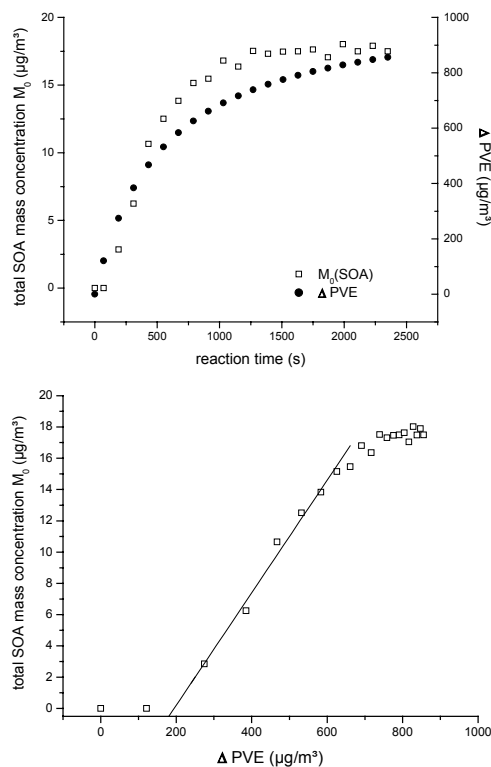


Fig. 4. Ozonolysis of PVE (initial reactant mixing ratios: 300 ppb): **(a)** total aerosol mass concentration formed and reacted PVE versus reaction time **(b)** total aerosol mass concentration formed versus reacted PVE.

5664

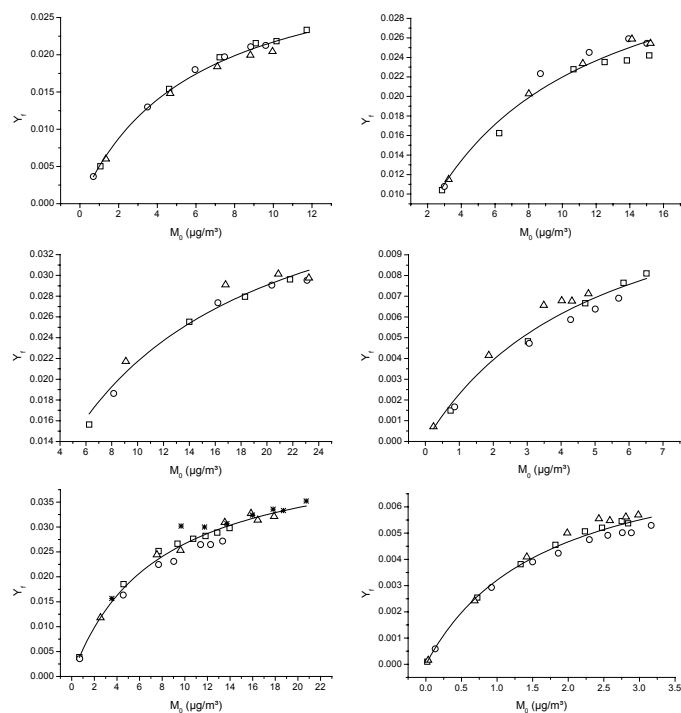


Fig. 5. SOA yield Y_f as a function of the aerosol mass M_0 during the ozonolysis of alkyl vinyl ethers (initial reactant mixing ratios: 300 ppb). Symbols (\square , \circ , Δ , $*$) designate measurements from different experiments under similar experimental conditions. **(a)** Ethyl vinyl ether **(b)** Propyl vinyl ether **(c)** *n*-Butyl vinyl ether **(d)** *t*-Butyl vinyl ether **(e)** *iso*-Butyl vinyl ether **(f)** *iso*-Butyl vinyl ether with an excess of cyclohexane.

5665

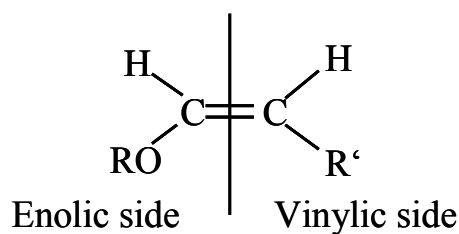


Fig. 6. Schematic structure of an enol ether.

5666

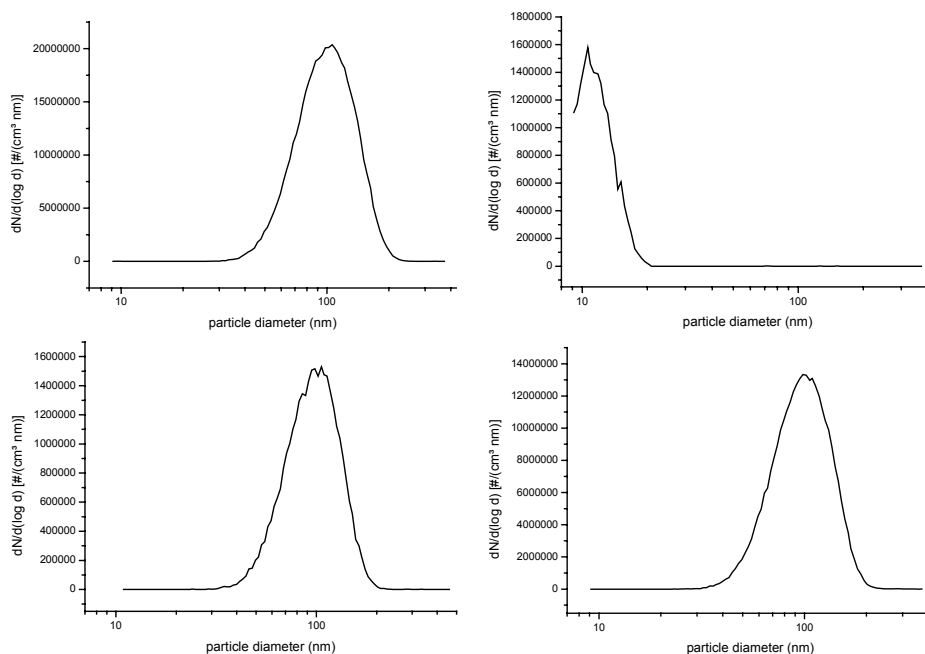


Fig. 7. SOA particle size distributions measured after completion of the reaction, just before the beginning of the filter sampling for chemical analysis by ESI(+)/MS. **(a)** EVE (9 ppm)+O₃ (8 ppm) (M_0 : 5 mg/m³) **(b)** EVE (9 ppm)+O₃ (8 ppm)+HCOOH (20 ppm) (M_0 <1 μg/m³) **(c)** EVE (9 ppm)+O₃ (8 ppm)+cyclohexane (300 ppm) (M_0 : 800 μg/m³) **(d)** EPE (9 ppm)+O₃ (8 ppm) (M_0 : 3 mg/m³).

5667

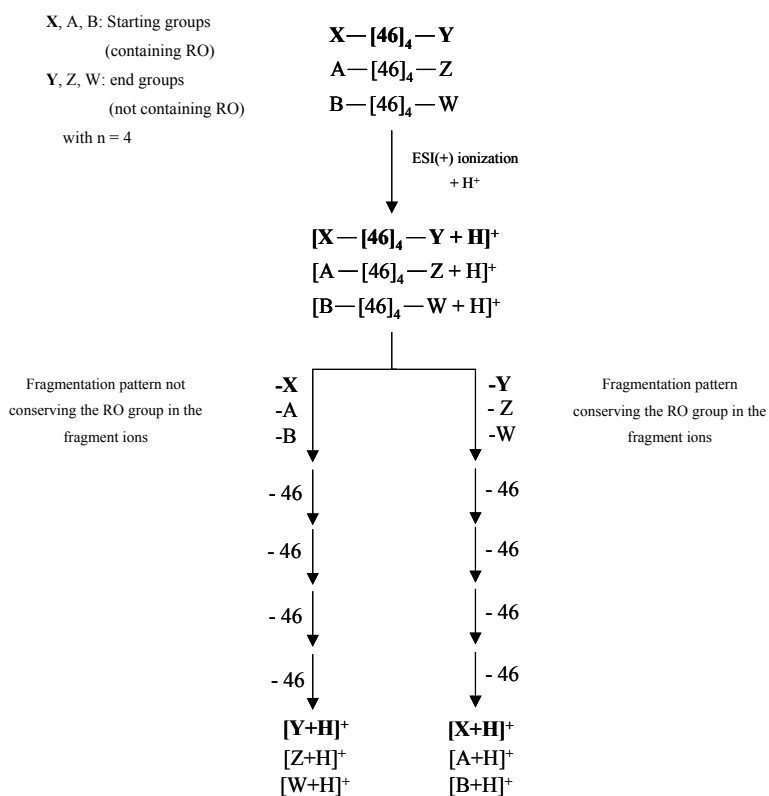
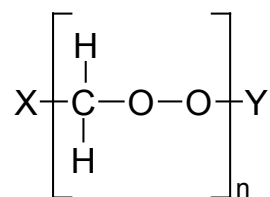
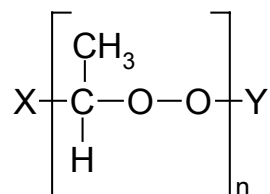


Fig. 8. SOA formed during ozonolysis of alkyl vinyl ethers AVE: Schematic diagram of the linear structure of the oligomers of type (a) and their MS/MS fragmentation involving the successive loss of the chain units of mass 46.

5668



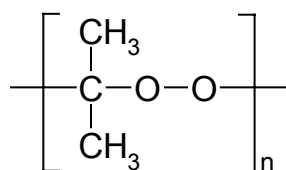
$\Delta m/z = 46$ (AVE)



$\Delta m/z = 60$ (EPE)

Fig. 9. Oligoperoxidic structure attributed to the oligomers formed during ozonolysis of **(a)** Alkyl vinyl ethers (AVE) **(b)** Ethyl propenyl ether (EPE).

5669



$\Delta m/z = 74$ (TME, liquid phase ozonolysis)

Fig. 10. Oligoperoxides formed during liquid phase ozonolysis of tetramethylethylene (TME) (Barton et al., 2004).

5670