

**Seasonal WSOC  
fraction of aerosol in  
the French Alps**

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# Seasonal variations of the Water Soluble Organic Carbon mass fraction of aerosol in two valleys of the French Alps

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## Abstract

Concentrations of Water Soluble Organic Carbon and WSOC fraction to OC were measured at two urban sites in valleys of the French Alps during a period of two and a half years. Concentrations were as high as 10–15  $\mu\text{g}/\text{m}^3$  in winter, but there is a clear seasonal cycle of the WSOC fraction, with minima occurring during winter. This reflects a marked dependency on temperature, with the average WSOC fraction being stable at  $54.8 \pm 7.7\%$  and  $75.9 \pm 6.3\%$  for temperatures in the ranges  $-10$  to  $+3^\circ\text{C}$  and  $12$  to  $24^\circ\text{C}$ , respectively. Several points are noteworthy in this evolution. First, there are limiting factors that prevent lower mass fractions in the low temperature range and higher mass fractions in the high temperature range. Second, the mass fraction at the lower temperature is rather high, in apparent contradiction with OC being mainly insoluble close to the emission sources. Third, the range of 20% for the change of the WSOC fraction between these extreme conditions is indeed rather narrow when compared to evaluations of the secondary (and supposedly water soluble) OC fraction proposed in the literature, with most of the published values being in the range 40 to 70%. A comparison of the evolution of WSOC concentrations with that of dicarboxylic acids (DCA) clearly indicates the influence of two regimes in the formation of WSOC: one at higher temperatures classically linked with the increase of DCA concentrations and associated with oxidation processes, and another at lower temperatures involving a much lower increase of DCA concentrations. We proposed several hypotheses involving processes that could be responsible for the large concentrations of WSOC in the particulate phase at our sites during winter.

## 1. Introduction

Few studies have focused on the Water Soluble Organic Carbon (WSOC) fraction of atmospheric particles, even though this fraction is probably of the utmost importance for aspects related to global climate (Kanakidou et al., 2005 and references therein) and

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the impact of aerosols on human health (Turpin, 1999; Dreher, 2000). Several investigations are currently available in the literature that outline studies conducted in urban or rural environments. However, the number of samples in most of these studies is rather low (see Sect. 3), and little ancillary data exists to delineate the respective impacts of sources and aging on the evolution of WSOC concentrations and the mass fraction relative to organic carbon (OC). Furthermore, these data were obtained through a large array of methods, both in relation to the collection of samples and the analyses of OC and WSOC concentrations, making comparisons and the construction of an overall picture rather difficult. The current belief maintains that the WSOC fraction is “low” for primary OC and that it increases with aging of the aerosol, together with the general oxidation state of organic matter (OM) (Saxena and Hildemann, 1996; Decesari et al., 2001). However, this idea has not been substantiated by many comprehensive field studies, and in particular, no full seasonal cycle of WSOC is currently published that presents data derived from homogeneous collection and analytical methods for various environmental conditions.

The program POVA (Pollution des Vallées Alpines) was launched in 2000 to focus on atmospheric chemistry in the two transit corridors between France and Italy: the Chamonix and Maurienne Valleys. This study was timed to take advantage of a unique opportunity: the “Tunnel du Mont Blanc” (TMB) in the Chamonix Valley was closed for nearly 3 years after a large accident in March 1999. During this time period, most of the heavy-duty traffic in the area was re-routed through the “Tunnel du Fréjus”, in the Maurienne Valley. The general goal of the POVA program was to compare the air quality and modelled atmospheric emissions and transport in these two valleys before and after the re-opening of the TMB to heavy-duty traffic. The program included several intensive field campaigns, a long-term campaign, and 3-D modelling of atmospheric dynamics and chemistry (Brulfert et al., 2005).

In this paper, we are presenting results of WSOC measurements conducted on samples obtained during the long-term sampling at one site in each valley. Other publications present discussions of the concentrations of EC and OC (Aymoz, 2005; Jaffrezo

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et al., 2005b<sup>1</sup>), and ionic species (Jaffrezo et al., 2005c<sup>2</sup>) obtained at the same time. We are aiming at an evaluation of the variability of the WSOC concentration and mass fraction in this study, and will make an attempt to understand the main factors and processes driving this variability. In a first part, we are presenting some tests performed to evaluate the experimental aspects of WSOC measurements.

## 2. Experimental

### 2.1. Sites and traffic

One site was located in the Chamonix valley (Fig. 1). This valley is 23 km in length, closed at its lower end by a narrow gorge (the Cluse pass) and at the upper end by the Col des Montets (1464 m a.s.l. (meters above sea level)) leading to Switzerland. The valley is rather narrow (1 to 2 km on average at the bottom). The elevation of the valley floor is approximately 1000 m a.s.l. on average, and is surrounded by tall mountains culminating in the summit of Mont Blanc (4807 m a.s.l.). There are no industrial emissions or waste incinerators in the valley, and the main anthropogenic sources of emissions are vehicular traffic, residential heating (typically fuel or wood-burning stoves), and some agricultural activity. The permanent population of about 12 000 is augmented by tourism, which brings in many times that number (on average 100 000 person/day in summer, and about 5 million overnight stays per year), mainly for short-term visits. There is only one road supporting all of the traffic into and out of the valley, but many secondary roads spread over the valley floor and the lower slopes. During the closing

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<sup>1</sup>Jaffrezo, J. L., Aymoz, G., Cozic, J., Chapuis, D., and Maenhaut, W.: Seasonal variation of PM10 main constituents in two valleys of the French Alps. I: EC/OC fractions, Atmos. Chem. Phys. Discuss., in preparation, 2005b.

<sup>2</sup>Jaffrezo, J. L., Aymoz, G., and Cozic, J.: Seasonal variation of PM10 main constituents in two valleys of the French Alps. II – Ionic constituents, Atmos. Chem. Phys. Discuss., in preparation, 2005c.

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of the TMB leading to Italy, the traffic at the entrance of the valley (14 400 vehicles/day on average) consisted mainly of cars (91% of the total, 50% of which were powered by diesel), along with a small number of local trucks (5%) and tourist buses (1%). Natural sources of emissions are limited to forested areas, which are mainly composed of coniferous species (95% of which consists of spruce, larch and fir). Re-suspension of soil dust is limited in winter, as most of the ground surface is covered by snow. The sampling site was located in the downtown area of the Chamonix town (6°52′16″ E; 45°55′26″ N; altitude 1038 m a.s.l.), about 10 m from a street. It can be considered a kerbside site.

The second site was located in the Maurienne valley in a residential urban area of Saint Jean de Maurienne (6°21′04″ E; 45°16′34″ N; altitude 555 m a.s.l.), about 2 km from both the highway and some large aluminium plants. It can be considered an urban background site. Saint Jean de Maurienne, the main town in the Maurienne Valley, is about 80 km from Chamonix in a SSW direction. This valley is much longer (about 80 km in length) and wider (about 3–5 km on average at the valley floor) than the Chamonix Valley. The surrounding mountains are lower, peaking at 3852 m a.s.l. with La Grande Casse. The elevation at the mouth of the valley is 330 m a.s.l. and increases to 2081 m a.s.l. at the end of the valley (the Col du Mont Cenis, leading to Italy). The topography of the Maurienne Valley is more complex than that of Chamonix, and includes many tributary valleys and several narrow gorges and sharp bends. The 30 km of the lower valley up to the town of Saint Jean de Maurienne supports several heavy industries, including steel transformation and aluminium and phosphorus production. Other anthropogenic emissions are linked to traffic. Most traffic (95% of the trucks and 55% of the cars) is found on the 50 km of highway leading from the valley entrance up to the town of Modane, the location of the entrance of the “Tunnel du Fréjus” leading to Italy. The traffic through the valley during the closure of the TMB averaged about 11 500 vehicles/day at Saint Jean de Maurienne, with 38% consisting of heavy-duty diesel trucks. The valley also supports agricultural activities, and some large forested areas, of which 60% consists of deciduous species (beech and chestnut) and the re-

maining 40% being occupied by coniferous trees (spruce). The total population of the valley is represented by about 45 000 inhabitants, most of whom are concentrated in the lower 40 km. Tourism is not as important as in Chamonix, but it results in about 4.2 million overnight stays per year on average, mostly during the winter season and in the upper part of the valley, where several ski resorts are located.

## 2.2. Sampling and processing

Both sites are permanent stations maintained by the local Air Quality Agency, l'Air de L'Ain et des Pays de Savoie. Continuous measurements at the stations conducted on a 15-min basis included NO<sub>x</sub> (Environnement SA, AC31M), ozone (Environnement SA, O341M), SO<sub>2</sub> (Environnement SA, AS21M), and PM<sub>10</sub> (TEOM1400, R and P). The TEOMs were heated at 50°C. Air intakes for all of these instruments were located about 4 m above ground on the roofs of the stations.

For this program, we added automatic aerosol sampling with ACCU samplers (R and P) on the by-pass flow of the TEOMs. These samplers were kept in a non-heated shelter close to outside temperature. All of the tubing of the ACCU systems was replaced with Teflon-coated lines. We used stainless-steel filter holders with a conical shape to insure homogeneous deposit on the filters. Sampling was conducted on a daily basis, from midnight to midnight (local time), at a flow rate of 0.85 m<sup>3</sup> hr<sup>-1</sup> (face velocity of 21 cm.s<sup>-1</sup>). The samplers were serviced once a week, with the collection of 7 samples and one field blank. We used QMA Whatman filters. These were pre-washed for 3 days in 3 successive bathes of Milli-Q water for lowering blanks for ionic species. After drying at 60°C, they were then fired for 1 h at 800°C to lower the blank levels for EC and OC. These filters were kept and transported to the field in packs of 20 in leak-proof glass jars before sampling. After sampling, exposed filters were downloaded directly in the field and stored at -5°C in individual petri slides wrapped in aluminium foil in sealed bags, until the analysis. All handling procedures were designed to reduce potential contaminations.

Sampling was conducted between 21 February 2001 and 3 July 2003 in Chamonix,

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and between 10 February 2001 and 25 June 2003 in Saint Jean de Maurienne. The number of samples and field blanks are presented in Table 1. While the data series are not totally continuous, the numbers of samples are sufficiently high to give a very good representation of the variability of the concentrations over the sampling periods.

### 5 2.3. Analysis for EC and OC

Samples were analyzed for EC and OC using the Thermo-Optical Transmission (TOT) method on a Sunset Lab analyzer (Birch and Cary, 1996). We used temperatures up to 870°C for the analysis of OC in 100% He, and up to 900°C for the analysis of EC in 98% He + 2% O<sub>2</sub> (Jaffrezo et al., 2005b<sup>2</sup>). Automatic split time was always used for the distinction between EC and pyrolysed OC, except for blank samples, where it was manually corrected when needed. The transit time between the front oven and the FID detector was optimized and adjusted at 13 s. We analyzed a fraction of 1.5 cm<sup>2</sup> from each sample, without any preparation. A subset of samples (approximately 5% of the overall set, for various sample loadings) was analyzed 3 times to check for the even repartition of particles on the filters and the accuracy of this sub-sampling method for the determination of atmospheric concentration. This test indicated a really good reproducibility of the deposition, with variability of results being within 3%.

The arithmetic average of the concentrations of the field blanks was taken into account for the calculation of atmospheric concentrations. Field-blank concentrations (equivalent to concentrations of OC=0.45±0.27 μg/m<sup>3</sup> and EC=0.045±0.064 μg/m<sup>3</sup> for a sample of 20 m<sup>3</sup> of air) account for 6.6 and 4.4% of the average atmospheric concentrations in summer for OC and EC, respectively (Aymoz, 2005). More than 760 daily concentrations for EC and OC were determined for each site (Table 1).

### 2.4. Extraction procedures for the determination of WSOC

We first tested two extraction procedures (soaking and filtration methods) aimed at removing the soluble fraction of OC from samples in order to subsequently analyze the

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remaining insoluble fraction on the filters with the TOT method. Some of the studies referenced in Table 3 use comparable methods. Our tests were performed on real samples collected on the roof of our institute (in Grenoble), in conditions comparable to those of actual POVA samples. The soaking method consists of the immersion of

5 the filter in  $n$  ( $n=1$  to 4) successive bathes of 10 ml of Elga<sup>®</sup> water for 10 min each. The filtration method consists of the filtration at a low flow rate (with the water sitting for 10 min prior the filtration) of  $n$  ( $n=1$  to 4) successive bathes of 10 ml of Elga<sup>®</sup> water through the sample held in a Millipore<sup>®</sup> filtration device. A final step of both methods included drying of the sample in a glove box in a nitrogen atmosphere before the

10 analysis, following the protocol mentioned above (Sect. 2.3.). The procedure was duplicated a number of times, depending on the number of water fractions and method. Figures 2 and 3 present the average percentage of recovery of EC and OC, respectively, in relation to the number of water fractions used for both methods. Results show that the samples are losing EC (about 22 and 48% by mass on average for the filtra-

15 tion and soaking methods, respectively) as a consequence of the first water fraction, with no further statistically significant changes in EC according to the number of water fractions. As EC is insoluble by definition, removal of these fractions most probably takes place by mechanical processes, with the likely candidates being resuspension in the case of soaking and entrainment in the case of washing. In these conditions, the

20 results obtained for OC (decreases of about 56 and 74% on average for the filtration and soaking methods, respectively) are difficult to interpret in terms of removal of the soluble fraction only. These methods were subsequently abandoned.

We finally used a method with the direct analysis of the soluble OC in the liquid phase (see below).  $N$  ( $n=1$  to 3, depending on the OC concentrations) aliquots (i.e. punches of 18 mm in diameter) of the initial daily samples were soaked in airtight glass bottles

25 in 17 ml of Elga<sup>®</sup> water for 40 min with some manual shaking. The extract was then filtered just before the analysis using Acrodisc filters (Pall Gelmann) with a porosity of 0.2  $\mu\text{m}$ .



## 2.5. Analysis of WSOC

The analyses of WSOC concentrations were performed with a Model 700 TOC analyzer (OI Analytical). It is based on the IR detection of the CO<sub>2</sub> formed after the hot chemical oxidation of the carbon content of the liquid sample. The sample is first acidified with 200 μl of 5% phosphoric acid to drive off, using a nitrogen flux for 2.5 min, the CO<sub>2</sub> formed from the inorganic carbonates. The OC content of the sample is then oxidized after the addition of 1 ml of sodium persulfate (100 g/l), and the CO<sub>2</sub> is subsequently measured after degassing with the nitrogen flux. The volume of the sampling loop was 10 ml, the oxidation time was set to 11 min, and the temperature for the oxidation was 100°C. These analytical conditions were previously tested by Pertuisot (1997).

The calibration was performed daily with 5 fresh synthetic solutions of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) covering the full range of the detector (0 to 4.2 μC/ml). This calibration was very stable from day to day. Several tests were performed to evaluate the uncertainty linked with this analysis. The coefficient of variation (standard deviation over mean) of 6 analyses of a synthetic solution at a concentration of 2.1 μC/ml was 0.7%. The average coefficient of variation for the analysis of real atmospheric samples (analyses of 3 punches for each of 3 actual samples) was 1.5%. This last value most probably represents the overall uncertainty of the analysis with this method.

A series of 45 blanks from the POVA campaign were analyzed for their WSOC content and were taken into account for the calculation of atmospheric concentrations. Overall, the average blank represents 7.2±1.9% of the average concentration in Saint Jean de Maurienne. This last series was analyzed totally, leading to 277 values of WSOC concentrations because of the accretion of consecutive samples (Table 1). Only 104 values were obtained for the Chamonix series, based on a selection of samples collected mainly during one summer and two winter periods.

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## 2.6. Analysis of organic dicarboxylic acids

The remaining parts of the filter (after extraction of the punches for EC/OC and WSOC analyses and removal of the outer ring in contact with the filter holder) were analyzed with Ionic Chromatography (IC) for the determination of a whole suite of anionic and cationic species (Jaffrezo et al., 2005c<sup>3</sup>) following the method described in Jaffrezo et al. (1998), Ricard et al. (2002) and Aymoz (2005). In brief, samples were soaked in 10 ml of Milli-Q water for 30 min in airtight glass bottles. They were then filtered just before the analysis using Acrodisc filters (Pall Gelmann) with a porosity of 0.2  $\mu\text{m}$ . Analysis of cations ( $\text{NH}_4^+$ ,  $\text{K}^+$ ,  $\text{Mg}^{2+}$ ,  $\text{Ca}^{2+}$ ) took place with a CS12 column on a Dionex 100 IC. The determination of  $\text{Na}^+$  was prevented by the high concentration remaining in the quartz filters. Analysis of inorganic anions ( $\text{Cl}^-$ ,  $\text{NO}_3^-$ ,  $\text{SO}_4^{2-}$ ) and dicarboxylic acids (DCA) (oxalate, glutarate, malonate, succinate, malate, and tartrate) took place on an AS11 column on a Dionex 500 analyzer. Table 2 presents the detection limits for these organic acids under the conditions of our analysis. We will only present results concerning the potassium concentration, as well as concentrations for the sum of all measured DCA. The full suite of the daily samples was analyzed for the series from Saint Jean de Maurienne, while only 536 samples were processed for the Chamonix series (Table 1).

## 2.7. Final data set

The final data set for this long-term sampling campaign also includes daily data for  $\text{NO}$ ,  $\text{NO}_x$ ,  $\text{O}_3$ ,  $\text{SO}_2$ , and  $\text{PM}_{10}$  concentrations, obtained by averaging the 15-min measurements. Furthermore, our investigation includes an analysis of meteorological data obtained from Météo France (temperature, precipitation, wind speed) from the stations in Chamonix and Saint Jean de Maurienne, with hourly measurements averaged to obtain daily values. We also obtained traffic data from regional administrations and highway companies for the sites of Tunnel des Chavants (at the lower entrance of the Chamonix valley) and Saint Michel de Maurienne (10 km upward of Saint Jean de Mau-

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rienne). These are daily data sets, with separate counts for individual cars and trucks. Finally, the data pool also includes daily concentrations for many VOC species, sampled in parallel to PM<sub>10</sub> with automatic systems during the first year of the campaign (Colomb, 2002). These measurements will not be discussed in this study.

### 3. Concentrations, mass fractions and seasonal variations of WSOC

Table 3 presents a comparison of concentrations and WSOC fractions derived from most of the studies available in the literature. With the exception of much higher concentrations in Tokyo (Sempere and Kawamura, 1994) and considerably lower concentrations at the high-altitude site of Jungfrauoch (Kryvacsy et al., 2001), all other concentrations (including our data) are in the same range. Our data indicates that concentrations of WSOC well above 10 μgCm<sup>-3</sup> can be measured in winter conditions in these urban areas, close to the emission sources (Fig. 4). This is much higher than concentrations obtained in “warm” conditions (T>15°C), which are generally considered favourable for the formation of WSOC. It should be noted that these data are, to the best of our knowledge, the first to include measurements at temperatures below freezing. We will not discuss the evolution of WSOC concentrations, which is strongly dependant upon atmospheric dynamics and local dispersion conditions. A presentation of factors influencing OC concentrations is presented elsewhere (Aymoz, 2005; Jaffrezo et al., 2005b<sup>2</sup>). However, potential sources of WSOC will be discussed below (see Sect. 4).

Average data for WSOC fractions range from a low of 12% (next to a highway in Paris; Ruellan and Cachier, 2001) up to 80% for a rural area in Sweden during summer (Zappoli et al., 1999) (Table 3). There is some tendency for lower values to be associated with urban areas, particularly in winter, and for higher values to occur in rural areas, especially in summer. This is more or less consistent with the hypothesis of WSOC consisting of a large fraction of secondary (oxidized) organic species. However, these data were obtained through very different protocols and are hardly comparable.

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No clear seasonal pattern has emerged so far from the only consistent data set derived from samples obtained in all seasons (Decesari et al., 2001). Our data largely confirms such a trend, with lower average values at both sites in winter compared to other seasons. Furthermore, these average values are nearly identical for the two sites during both summer and winter, despite their slightly different characteristics (kerbside for Chamonix, urban background for Saint Jean).

Figure 5 presents all measurements of the WSOC fraction obtained in this study. It shows a clear seasonal cycle with minima during both winters (2002 and 2003) for the two sites. While the values during the 3 summer periods are rather similar (averages between 75–80%), the minimum values during the winters are different, reaching (for both sites) 40–45% for the winter of 2002, but remaining above 50% during the winter of 2003. This suggests a strong correlation with changes in temperature.

All of these observations are confirmed in Fig. 6, which shows the evolution of the WSOC fraction in relation to temperature for both sites, together with the moving average (and standard deviation) for each 2°C increment. A clear pattern of change emerges, with the average WSOC fraction being stable at  $54.8 \pm 7.7\%$  and  $75.9 \pm 6.3\%$  for temperatures below 3°C (in the range  $-10$  to  $+3^\circ\text{C}$ ) and above 12°C (in the range 12 to 24°C), respectively. There is a constant increase of the WSOC fraction between these ranges. These observations indicate that, on average, there are processes that prevent on the one hand WSOC fractions higher than 80%, even for so-called favourable oxidation conditions, and on the other hand WSOC fractions lower than 40% for conditions that do not favour the oxidation of primary OC.

However, dispersion is still rather high (average coefficient of variation about 11.1% for all of the 2°C intervals), and indicates that temperature (taken as a gauge of oxidation conditions) is not the only factor affecting observed patterns. It will therefore be interesting to relate these changes with other indicators of oxidation or tracers of sources.

#### 4. Relation between WSOC, ozone, organic diacids and potassium

The evolution of the WSOC fraction has already been positively related to ozone concentrations in some studies (Sullivan et al., 2004), indicating a degree of dependency between processes that lead to the formation of oxidized species in the gas and particulate phases. However, it has been shown that such a relation does not hold in all environmental conditions, with no further gain in the understanding of the connection between the processes (Sullivan et al., 2004). In our case, there are also some degrees of co-variation between ozone concentrations and the WSOC fractions at both sites, but the relations are not strong:

$$\begin{aligned} \text{WSOC fraction (\%)} &= 0.28 * [\text{ozone } (\mu\text{gm}^{-3})] + 56.9; r^2=0.45; n=277 \text{ in Saint Jean} \\ \text{WSOC fraction (\%)} &= 0.32 * [\text{ozone } (\mu\text{gm}^{-3})] + 51.3; r^2=0.51; n=104 \text{ in Chamonix.} \end{aligned}$$

However, the correlations between WSOC and ozone concentrations are much lower:

$$\begin{aligned} \text{WSOC } (\mu\text{gm}^{-3}) &= -0.026 * [\text{ozone } (\mu\text{gm}^{-3})] + 5.81; r^2=0.18; n=277 \text{ in Saint Jean} \\ \text{WSOC } (\mu\text{gm}^{-3}) &= -0.065 * [\text{ozone } (\mu\text{gm}^{-3})] + 9.65; r^2=0.29; n=104 \text{ in Chamonix.} \end{aligned}$$

The absence of a correlation can reflect the influence of many processes, such as differences in oxidation and removal processes, their time scales, temperature dependence, or differences in the concentrations of primary species. Indeed, a more interesting analytical approach concerns the comparison of concentrations of WSOC with concentrations of species of known secondary origin in the particulate phase. This is the case for dicarboxylic acids (DCA), which are generally regarded as end products of the oxidation of primary species, even if primary sources appear to be present in the urban atmosphere (Chebbi and Carlier, 1996). The most likely oxidation pathways for the production of DCA are probably in aqueous phase (hydrated particles

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or cloud droplets) after an initial gas phase oxidation of primary compounds (Ervens et al., 2004), since direct production in the gas phase seems unlikely (Aumont et al., 2000).

Figure 7 presents the change of concentrations of the sum of DCA measured in our samples in relation to temperature. It shows a large continuous increase of the concentrations for warmer temperatures, and is consistent with a secondary production by oxidation. The concentrations do not reach a plateau for the warmer temperatures, an indication that the production of these species with low vapour pressures is not the limiting factor explaining the maximum ratio at 80% for the WSOC fraction (Fig. 6). The concentrations are still significant for cooler temperatures, possibly indicating primary production or a degree of secondary formation. Indeed, this behaviour at sub-freezing temperatures is not the same for all species investigated (Aymoz, 2005), with some (tartrate, malate) showing near-zero concentrations while others (oxalate) exhibit significant concentrations. This finding will not be discussed further in this paper.

Figure 8 presents the mass fraction of WSOC represented by the measured DCA in relation to temperature. Again, we observe a large and continuous increase with increasing temperature which does not show any sign of reaching a limitation. This mass fraction reaches a level of about 10–14%, and is in the range discussed by Saxena and Hildemann (1996) for observations made in urban sites. This mass fraction is much lower at low temperatures, and measures of DCA never account for more than 3% of the WSOC concentrations for  $T_{\text{OC}} < 4^{\circ}\text{C}$ .

Indeed, Figs. 9 and 10 show that very different responses exist for WSOC and DCA on each side of a threshold temperature, with large increases of DCA concentrations occurring with moderate increases for WSOC concentrations for the higher temperatures. Conversely, the increase of DCA is only minor when large increases in WSOC concentrations occur at the lower temperatures. This clearly indicates two regimes for the formation of WSOC: one being closely related to oxidation processes producing DCA (at warmer temperatures), and another (at lower temperatures) being more loosely dependent upon the formation of these secondary species.

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The slopes and intercepts of the regressions (Figs. 9 and 10) for the two sites are similar in both regimes, and indicate that the processes may be the same in the two valleys. However, a difference of 4°C can be seen in the threshold temperature between the two sites, corresponding to the difference in the average daily temperature at the sites:

$$T^{\circ}\text{C (CHX)}=0.99 * T^{\circ}\text{C (STJ)} - 4.3 \quad (r^2=0.96; n=2145; 18 \text{ July } 1997 \text{ to } 1 \text{ June } 2003).$$

A likely hypothesis to explain this difference is that the link between WSOC and DCA for the warmer temperatures results from oxidation reactions that take place at similar temperatures at places further away from the sampling sites, with the aerosols being subsequently transported into the valleys (with the temperature at each receptor site being dependent on altitude, among others factors). The impact of imported air-masses in the valleys in summer has already been demonstrated in the case of ozone (Brulfert, 2004; Brulfert et al., 2005). Furthermore, this hypothesis could explain the same shift of 4°C that is also seen in Figs. 4 and 7 for the increases in the concentrations of WSOC and DCA in the warmer temperature ranges. This implies that a large share of the summer PM<sub>10</sub> particles (at least for the secondary species) at both sites is coming for transport, and that it is not produced locally. Such a shift in the case of WSOC probably implies that most of the WSOC results from oxidation processes with a characteristic time dependent on the transport time to the sites, with a small share of soluble primary compounds or secondary species produced locally, during the summer periods.

The case involving cold temperatures is more complex, exhibiting a much weaker correlation when considering the combined data set:

$$\begin{aligned} \text{DCA } (\mu\text{g}/\text{m}^3) &= 0.124 * \text{WSOC } (\mu\text{g}/\text{m}^3) - 0.28 \quad (r^2=0.54) \text{ for warm conditions } (n=206) \\ \text{DCA } (\mu\text{g}/\text{m}^3) &= 0.011 * \text{WSOC } (\mu\text{g}/\text{m}^3) + 0.06 \quad (r^2=0.26) \text{ for cold conditions } (n=173). \end{aligned}$$

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This may reflect the impact of several processes, or of different processes in each valley, on the production of WSOC in winter. It should be remembered that the characteristics of aerosol at our two sites are more strongly influenced by local conditions (primary emissions, meteorology) in winter than in summer periods (Aymoz, 2005). At least three non-exclusive hypotheses can be proposed to explain both the weak link between DCA and WSOC, and the large WSOC fraction during winter: i) direct emissions of soluble species, ii) oxidation pathways differ between summer and winter periods, and iii) changes in the gas/particle repartition for semi-volatile species.

The strong influence of temperature on the repartition of semi-volatile species between gas and particulate phases (described by the coefficient  $K_p$ ) is well known (Pankow, 1994), and was verified in our case for some PAH species (Marchand, 2003). It has been shown, for example, that  $K_p$  can change by nearly 2 orders of magnitude between 21°C and -3°C for some semi-volatile species (Leaitch et al., 1999), a change that is dependent upon the enthalpy of vaporization of the species (Tsigaridis and Kanakidou, 2003; Pun et al., 2003). Such a response has been reported for soluble aldehydes (Saxena and Hildemann, 1996). Therefore, larger concentrations of such species (both primary and secondary) could be transferred in the particulate phase during winter. Additionally, partitioning for each species is dependent upon the composition of the aerosol that is already condensed, with  $K_p$  greatly increasing with the affinity between the condensed and condensable products (Leaitch et al., 1999; Chandramouli et al., 2003). Residential biomass burning is a significant source of OC in winter in the valley (see below), with a large share of the emission products being composed of polar compounds, potentially increasing the condensation of this class of soluble species. However, detailed studies of the relation between the particulate and gaseous phases together with further molecular speciation of OC and WSOC are needed to verify if such a process can have a significant impact on the WSOC fraction of OC at low temperatures.

The oxidation of primary species during winter conditions is also possible, but it is still rather difficult to evaluate its importance in terms of the chemical evolution of



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OC. Measurements of the size distribution of OC in winter during POVA (Jaffrezo and Aymoz, 2005) show that a very large fraction of its mass ( $61 \pm 2\%$ ,  $n=6$  daily samples collected in January 2003) exists in the so-called droplet mode (aerodynamic diameter in the range 0.26 to  $1.0 \mu\text{m}$ ), which is generally associated with cloud processing of smaller particles (Blando and Turpin, 2000; Ervens et al., 2004). A much smaller fraction of the OC mass exists in the lower size ranges ( $15 \pm 1\%$ , for the same samples) associated with primary emissions, such as those produced from diesel (Kerminen et al., 1997) or wood-smoke emissions (Hays et al., 2004). Oxalate is also measured in the same samples and exists mainly in this droplet mode, albeit with a much lower mass ratio to OC than that found in summer samples. It is likely that a large share of this cloud processing during this period is taking place locally in the fog that formed at the top of the low inversion layers during these anti-cyclonic conditions. The processes involved in the formation of such a droplet mode at low temperature could lead to repartition of soluble organic species that differ from those of the summer period, and may potentially include aldehydes and other compounds (Blando and Turpin, 2000), as well as polyacidic compounds (Decesari et al., 2001). Even if polyacidic compounds were present in larger proportions in the winter rather than summer samples in the study of Decesari et al. (2001), this hypothesis needs to be tested extensively and compared to the addition of OC mass in the particulate phase via a polymerization mechanism as proposed by Kalberer et al. (2004).

Finally, we can try to test the hypothesis concerning a change in the sources of OC that would increase WSOC concentrations in winter independently of DCA production. We have several indications that residential wood burning represents a substantial share of the emissions of OC in winter in the valleys (Marchand et al., 2004, 2005<sup>3</sup>; Aymoz, 2005), and it is well known that these emissions include a large proportion of soluble organic species (Mayol-Bracero et al., 2002; Schauer et al., 2001). We can at-

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<sup>3</sup>Marchand, N., Aymoz, G., Jaffrezo, J. L., Masclet, P., and Besombes, J. L.: Biomass burning indicators in the aerosol of Alpine valleys, Atmos. Chem. Phys. Discuss., in preparation, 2005.

tempt an evaluation of the impact of this source between cold and warm conditions by considering the ratio with soluble potassium, a species also emitted by wood burning. Figure 11 shows the correlation between OC and  $K^+$  using the overall data set for the two sites.

5 Excluding a few points (13 and 14 July 2001, 31 December 2001, and 1 January 2003) that are influenced by fireworks, the correlation for the Chamonix series is high (Fig. 11), while the correlation is lower for samples from the Maurienne Valley. The very same correlations are observed when considering WSOC and  $K^+$ :

10 Chamonix:  $WSOC (\mu gC/m^3) = 26.7 * K^+ (\mu g/m^3) + 3.5 (r^2=0.82; n=95)$

Saint Jean:  $WSOC (\mu gC/m^3) = 6.8 * K^+ (\mu g/m^3) + 3.7 (r^2=0.62; n=277).$

The  $OC/K^+$  ratios shown in Table 4 can be compared to the values of 71–85 presented by Khalil and Rasmussen (2003) for measurements in Olympia (WA, USA) under conditions strongly influenced by residential wood burning, or to measurements of direct fireplace emissions performed by Schauer et al. (2001) that yielded ratios of 202 (for pine) and 91 (for oak) for the fine fraction (diameter  $<1.8 \mu m$ ) of the aerosol. Both studies measured elemental K. The ratio observed for Saint Jean in our study most probably indicates additional sources of potassium. Taking into account all of the uncertainties associated with the measurements of OC, both the ratio  $OC/K^+$  in Chamonix and the high correlation between OC and  $K^+$  are more suggestive of a significant impact by residential wood burning. It is therefore possible that this emission source influences WSOC winter concentrations in Chamonix. However, the small changes of the  $OC/K^+$  ratio between winter and summer periods remain unclear, and further investigations are needed to substantiate a hypothesis of a WSOC fraction largely modulated by residential wood-burning emissions in both valleys during winter.

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## 5. Conclusions

Concentrations of Water Soluble Organic Carbon and WSOC fraction to OC were measured at two urban sites in valleys of the French Alps during a period of two and a half years. Concentrations were as high as 10–15  $\mu\text{g}/\text{m}^3$  in winter, but there is a clear seasonal cycle of the WSOC fraction, with minima occurring during winter. This reflects a marked dependency on temperature, with the average WSOC fraction being stable at  $54.8 \pm 7.7\%$  and  $75.9 \pm 6.3\%$  for temperatures in the ranges  $-10$  to  $+3^\circ\text{C}$  and  $12$  to  $24^\circ\text{C}$ , respectively.

Several points are noteworthy in this evolution. First, there are limiting factors that prevent lower mass fractions in the low temperature range and higher mass fractions in the high temperature range. Second, the mass fraction at the lower temperature is rather high, in apparent contradiction with OC being mainly insoluble close to the emission sources. Third, the range of 20% for the change of the WSOC fraction between these extreme conditions is indeed rather narrow when compared to evaluations of the secondary (and supposedly water soluble) OC fraction proposed in the literature, with most of the published values being in the range 40 to 70% (Castro et al., 1999; Strader et al., 1999; Na et al., 2004). All of these points deserve further investigation, both in other environmental conditions and in association with other measurements (molecular speciation, evaluation of chemical functional groups, etc.).

A comparison of the evolution of WSOC concentrations with that of dicarboxylic acids (DCA) clearly indicates the influence of two regimes in the formation of WSOC: one at higher temperatures classically linked with the increase of DCA concentrations and associated with oxidation processes, and another at lower temperatures involving a much lower increase of DCA concentrations. We proposed several hypotheses involving processes that could be responsible for the large concentrations of WSOC in the particulate phase at our sites during winter. None of these hypotheses can be validated or dismissed with our supporting data. Investigations on the WSOC fraction at low temperatures are needed, for such a temperature range is of the utmost importance

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for cloud formation.

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**Table 1.** Sampling dates, and statistics on sampling at both sites.

	Chamonix	Saint Jean de Maurienne
Date start	21 Feb 01	10 Feb 01
Date stop	03 Jul 03	25 June 03
Number of days	862	864
Number of valid samples	785	798
% of valid samples	91.1	92.4
Number of field blanks	108	113
Number of EC/OC data	763	788
Number of WSOC data	104	277
Number of OA and K <sup>+</sup> data	537	777

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**Table 2.** Average detection limits and fraction of samples below detection limit for potassium and the dicarboxylic acids measured in this study.

		Detection limit ( $\text{ng/m}^{-3}$ )	% sples <dl (Chamonix)	% sples <dl (Saint Jean)
Potassium	K+	12.0	6	0
Oxalate	$\text{C}_2\text{O}_4^{2-}$	2.4	0	0
Malonate	$\text{CH}_2\text{C}_2\text{O}_4^{2-}$	<1	22	14
Succinate	$(\text{CH}_2)_2\text{C}_2\text{O}_4^{2-}$	<1	0	0
Glutarate	$(\text{CH}_2)_3\text{C}_2\text{O}_4^{2-}$	<1	8	13
Malate	$\text{CH}_2\text{CH}(\text{OH})\text{C}_2\text{O}_4^{2-}$	<1	14	6
Tartarate	$(\text{CH}(\text{OH}))_2\text{C}_2\text{O}_4^{2-}$	<1	6	10

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**Table 3.** WSOC concentrations and mass fractions in the literature. N: number of samples. WSOC fraction of OC, except when noted. References 1: Sempere and Kawamura (1994), 2: Zappoli et al. (1999), 3: Decesari et al. (2001), 4: Krivacsy et al. (2001), 5: Kiss et al. (2002), 6: Sullivan et al. (2004), 7: Ruellan and Cachier (2001), 8: this study, with winter as DJF, etc., \*: continuous measurements.

Site	Dates	N	WSOC ( $\mu\text{g}/\text{m}^3$ )	WSOC fraction (%)	Ref
Tokyo (urban)	Winter 92	22	3.2–3.4	28–32% of TC	1
	Summer 92		21.3–23.2	32–55% of TC	
Sweden (rural)	June/July 96	5	1.7	77	2
K-Pusztza (rural)	July/August 96	5	2.4	48	
Italy (urban)	September 96	6	4.0	65	
Pô Valley (rural)	Jan/Feb 98	9	7.1	47	3
	Mar/Apr 98	10	2.3	49	
	May/Sep 98	9	1.4	50	
	Oct/Nov 98	4	3.4	47	
	Oct/Feb 98–99	4	5.7	38	
	Mar/Avr 99	4	2.4	38	
Jungfrauoch (altitude)	Summer 98	8	0.63	60	4
K-Pusztza (rural)	Summer 98	22	4.8	63	
K-Pusztza (rural)	Jan–Sept 00	42	2.0–8.25	66% of TC	5
St Louis (urban)	June 03	*	2.87	64	6
	Aug 03		2.40	61	
	Oct 03		1.33	31	
Paris (kerb side)	Aug–Oct. 97	4	4.4	12.4	7
Chamonix (urban)	Winter	69	8.8±3.1	54.5±6.0	8
	Summer	26	5.6±1.0	75.9±5.9	
Saint Jean (urban bckgrnd)	Winter	78	5.8±2.4	56.6±10.0	
	Spring	93	3.9±0.8	76.0±5.7	
	Summer	42	4.6±1.0	77.0±8.5	
	Fall	64	4.7±1.6	68.8±9.8	

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**Table 4.** Values of the ratio  $OC/K^+$  at the two sites.

	OC/K <sup>+</sup> mean	Standard deviation	Median	N
Chamonix				
T<3°C	108.3	43.2	94.7	202
T>3°C	134.3	62.4	119.5	318
Saint Jean				
T<7°C	48.1	25.2	41.6	246
T>7°C	51.8	13.7	47.3	526

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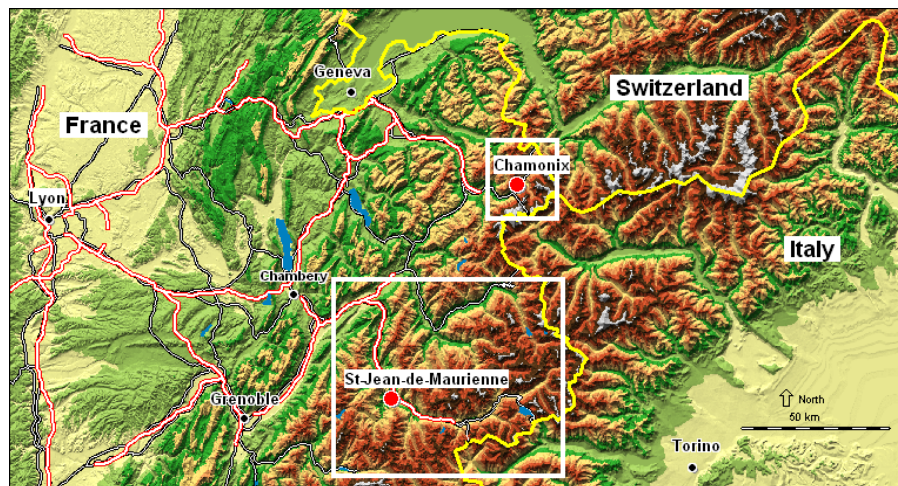
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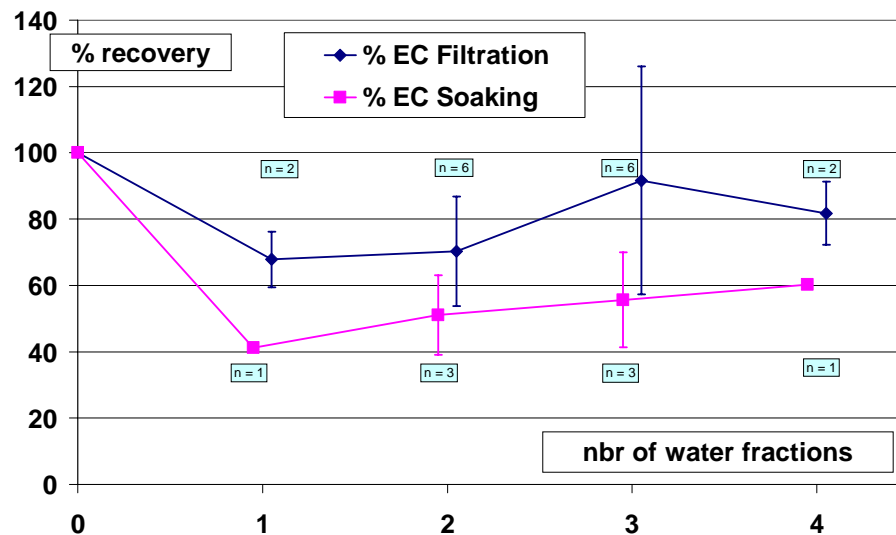
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**Fig. 1.** Sampling areas in the French Alps.[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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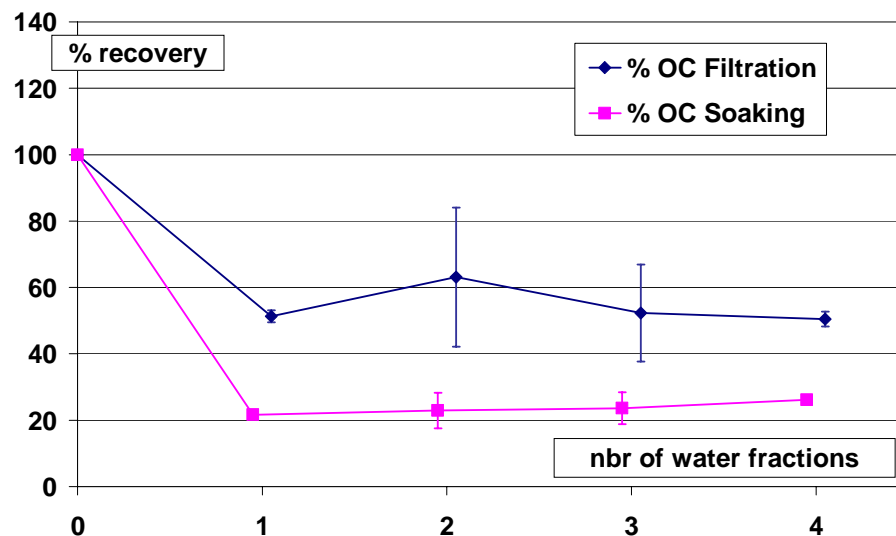
**Fig. 2.** Recovery of EC after washing with filtration and soaking methods, according to the number of water fractions (10 ml each) used for the washing. *n* is the number of duplicates and the error bar for each point represents  $\pm 1$  standard deviation of the measurements.

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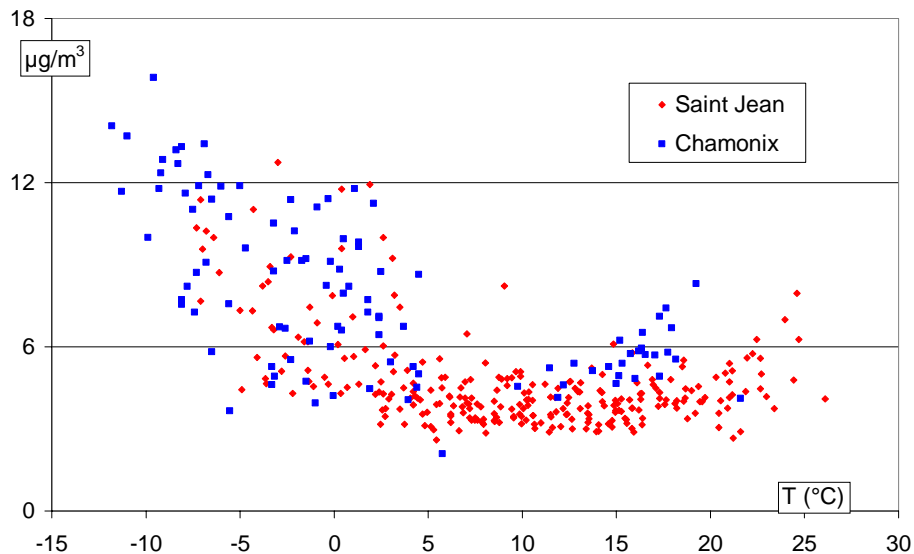
**Fig. 3.** Recovery of OC after washing with filtration and soaking methods, according to the number of water fractions (10 ml each) used for the washing. The number of duplicates is the same as for EC (Fig. 2) and the error bar for each point represents  $\pm 1$  standard deviation of the measurements.

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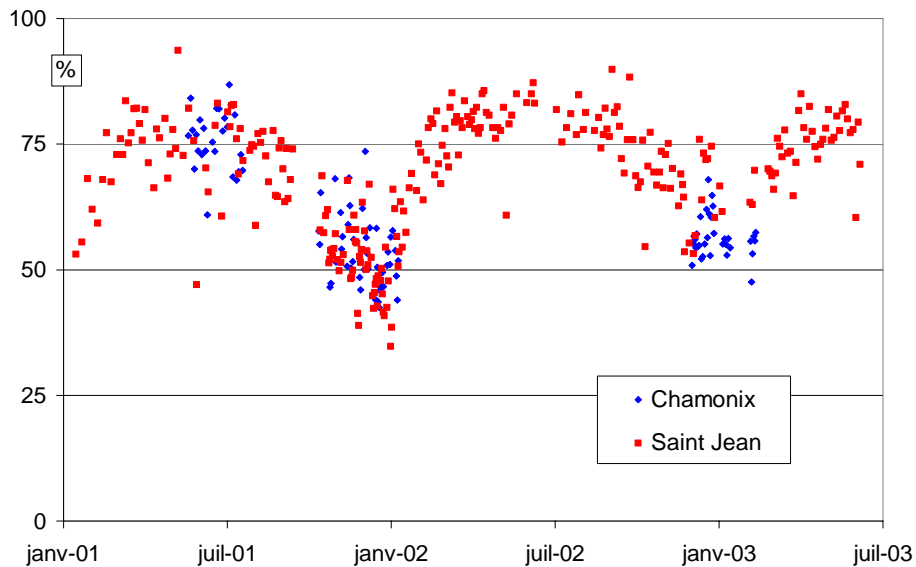
**Fig. 4.** Evolution of daily WSOC concentrations according to the temperature for the two sites.

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**Fig. 5.** Evolution of the WSOC fraction for the overall sampling period at the two sites.

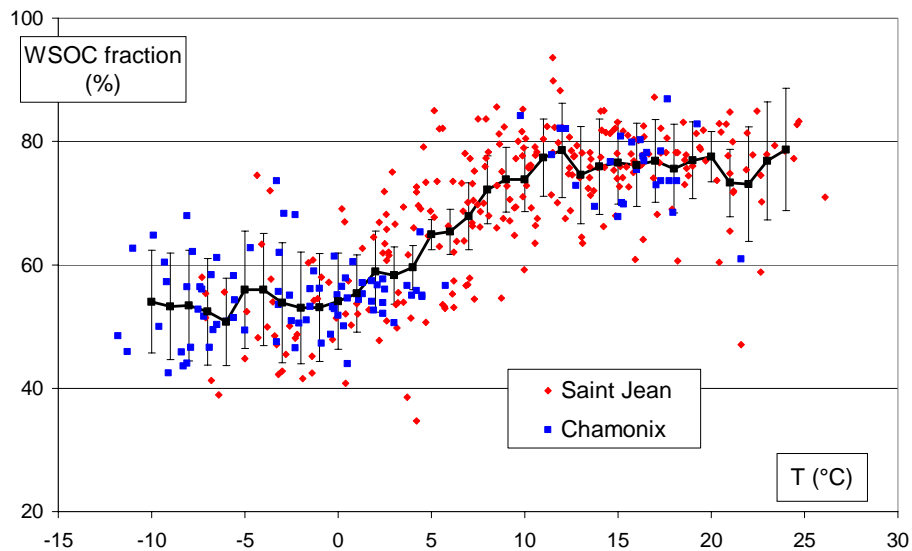
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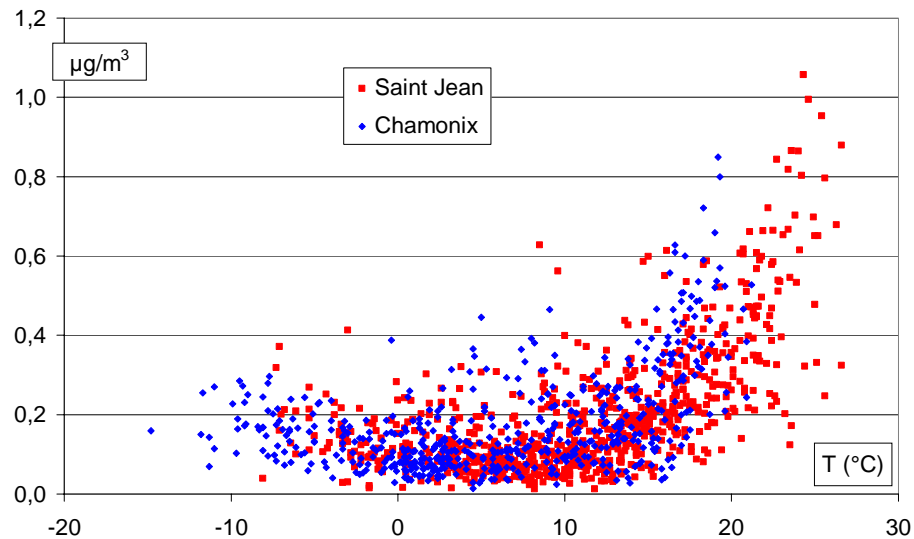
**Fig. 6.** Evolution of the WSOC fractions according to the temperature for the two sites.

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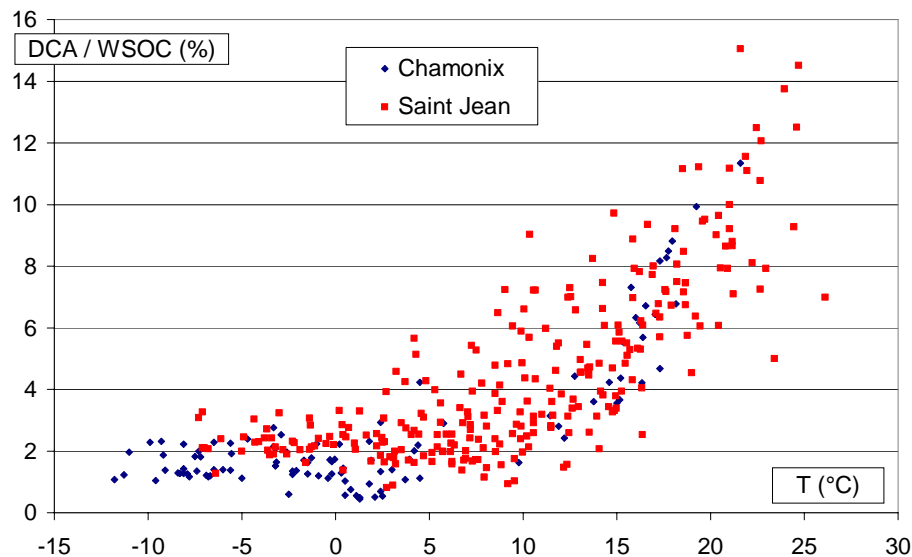
**Fig. 7.** Evolution of the sum of the concentrations of the dicarboxylic acids (DCA) measured, according to the temperature for the two sites.

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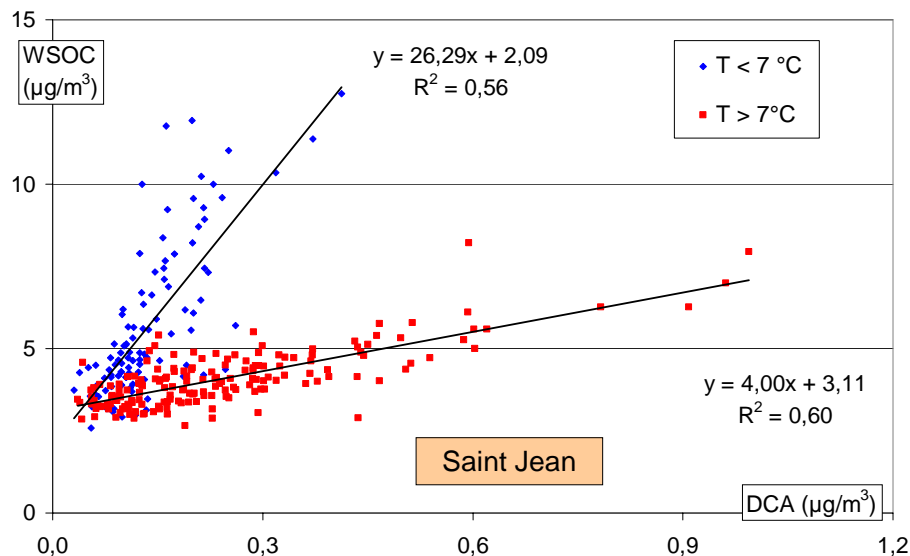
**Fig. 8.** DCA fraction of the WSOC according to the temperature at the two sites.

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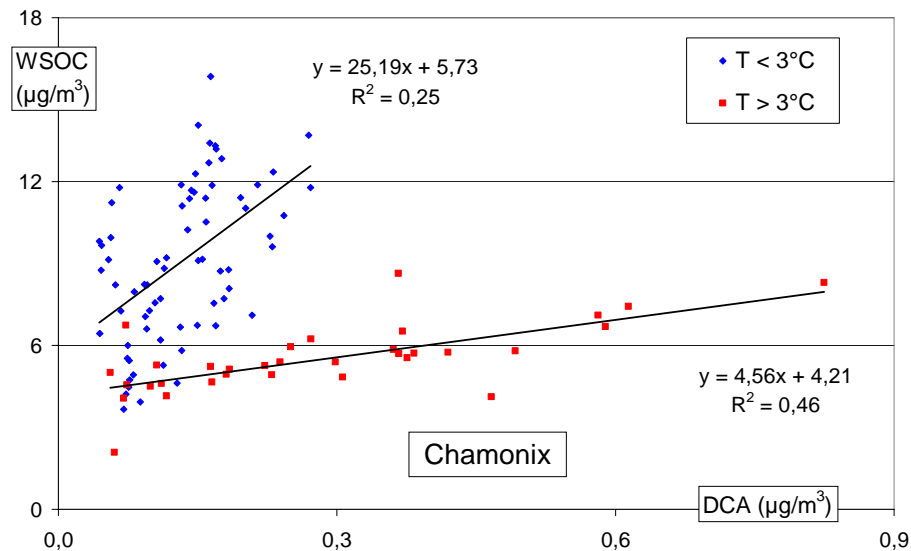
**Fig. 9.** Dicarboxylic acids and WSOC concentrations according to the temperature in Saint Jean de Maurienne.

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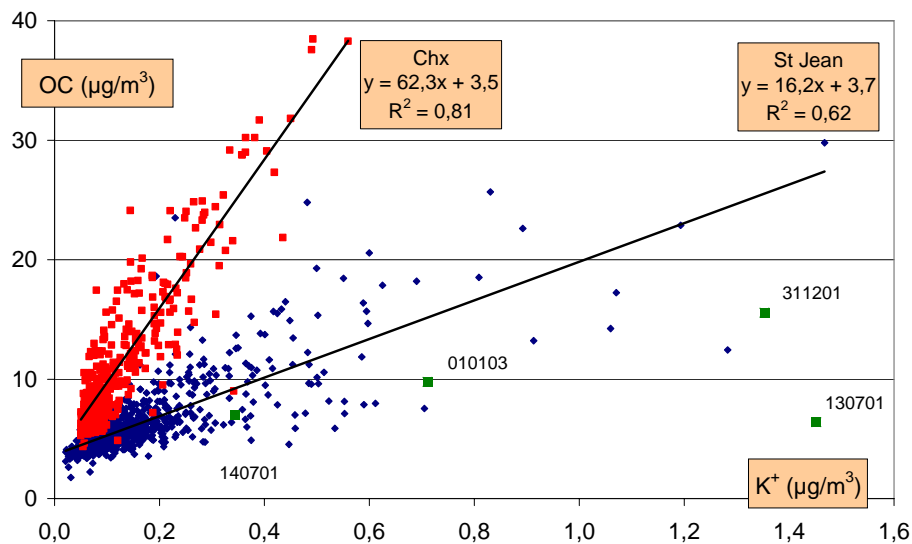
**Fig. 10.** Dicarboxylic acids and WSOC concentrations according to the temperature in Chamonix.

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**Fig. 11.** Potassium and OC concentrations at both sites.[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

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