

**Intercomparison of  
two isoprene  
measurement  
techniques**

Leibrock et al.

# Ground-based intercomparison of two isoprene measurement techniques

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Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

## Abstract

An informal intercomparison of two isoprene ( $C_5H_8$ ) measurement techniques was carried out during Fall of 1998 at a field site located approximately 3 km west of Boulder, Colorado, USA. A new chemical ionization mass spectrometric technique (CIMS) was compared to a well-established gas chromatographic technique (GC). The CIMS technique utilized benzene cation chemistry to ionize isoprene. The isoprene levels measured by the CIMS were often larger than those obtained with the GC. The results indicate that the CIMS technique suffered from an interference associated with air masses from the Denver, CO metropolitan area. However, the CIMS technique is also demonstrated to be sensitive and fast and is a candidate for isoprene measurements in remote environments near isoprene sources.

## 1. Introduction

Isoprene is a  $C_5H_8$  hydrocarbon that is emitted in large quantities by deciduous trees and other plant species, (e.g. [Singh and Zimmerman, 1992](#); [Steinbrecher, 1997](#)). Isoprene reacts at the gas kinetic rate with the OH radical and has an atmospheric lifetime of approximately two hours during the day. Consequently, isoprene plays an important role in the production of secondary pollutants such as ozone, PAN, and CO ([Atkinson, 1990](#); [Montzka et al., 1993](#)). Isoprene mixing ratios in the lower atmosphere vary from a few pptv to several ppbv and depend on factors such as land use, season and temperature. Isoprene also reacts with ozone, which provides an important loss mechanism during the night. A detailed review of the atmospheric chemistry of isoprene as well as estimates of the global emissions is given in [Fehsenfeld et al. \(1992\)](#).

In order to better understand oxidant formation on local and regional scales, reliable airborne measurements of the temporal and spatial distribution of isoprene are needed, requiring sensitive, fast time response ( $t \sim 1$  s) techniques. We developed a chemical ionization mass spectrometric method (CIMS) for the detection of isoprene in

## Intercomparison of two isoprene measurement techniques

Leibrock et al.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

---

**Intercomparison of  
two isoprene  
measurement  
techniques**Leibrock et al.

---

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[⏪](#)[⏩](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

ambient air (Leibrock and Huey, 2000). CIMS techniques are generally known for their high time resolution and good sensitivity and have been applied to measure a number of different trace gases, including nitric and formic acid (Huey and Lovejoy, 1996), OH (Eisele et al., 1994), ammonia, acetonitrile, acetone (Arnold and Hauck, 1985), and various other organic compounds (Hansel et al., 1995; Lindinger and Hansel, 1997). The goal of this study is to characterize the CIMS isoprene measurement by comparison with a well-established gas chromatographic (GC) technique (Goldan et al., 2000).

## 2. Experimental section

### 2.1. Intercomparison location and time frame

Isoprene measurements were performed on 18, 29, 30 September and 1 October 1998 at the Enchanted Mesa field site located ~ 3 km west of Boulder, Colorado at 1770 m above mean sea level in the foothills of the Rocky Mountains at the northwest edge of the Denver metropolitan area. At this location mixing ratios of VOCs such as isoprene can vary by orders of magnitude within a short time frame. These variations are primarily due to wind speed and direction and to a lesser extent to other factors such as temperature and precipitation. When the wind is from the west, clean air with low isoprene mixing ratios is transported to the site from the mountains. When the wind is from the urban areas to the south or east, large VOC loadings and high levels of other pollutants such as nitrogen oxides can be transported to the site. Consequently, the site is well suited for a challenging instrument intercomparison and has been used for this purpose on previous occasions (Fehsenfeld et al., 1998; Williams et al., 1992). Ancillary measurements of a wide range of chemical species and meteorological parameters were also performed during the intercomparison.

## 2.2. CIMS

The CIMS instrument was located in a trailer on the mesa. The air sampling inlet extended horizontally  $\sim 50$  cm from the west wall of the trailer at a height of  $\sim 2$  m above the ground. The inlet consisted of a 1 m long  $\times$  2.5 cm OD Teflon tube in which a flow of 17.5 SLPM (standard liters per minute) was maintained with a carbon vane pump. A small fraction of the total inlet flow,  $\sim 3$  SLPM, was sampled into the CIMS. The CIMS consisted of a flow tube reactor that was coupled to a quadrupole mass spectrometer (Fig. 1). The flow tube was equipped with a radioactive ion source ( $^{210}\text{Po}$ , NRD) and a 1 mm diameter air sampling orifice. A total pressure of 20 Torr was maintained in the flow tube by a rotary vane pump. Benzene monomer and dimer cations,  $\text{C}_6\text{H}_6^+$  and  $(\text{C}_6\text{H}_6)_2^+$ , were synthesized by flowing  $\sim 3$  SLPM of nitrogen doped with  $\sim 10$  ppmv benzene vapor through the ion source into the flow tube reactor. The constituents of the sampled air reacted with the benzene cations along the length of the flow tube (reaction time  $\sim 100$  ms). The resulting ions were sampled into the mass spectrometer where they were mass filtered and detected with an ion multiplier. The pulses from individual ions were amplified, counted and stored in a computer.

Isoprene reacted by charge transfer association reaction with the benzene reagent ions (Leibrock and Huey, 2000):



Both benzene reagent ions reacted with isoprene to form the same product ion,  $\text{C}_5\text{H}_8^+(\text{C}_6\text{H}_6)$ , at 146 amu. Isoprene was detected by sequentially monitoring masses 78 amu ( $\text{C}_6\text{H}_6^+$ ), 156 amu ( $(\text{C}_6\text{H}_6)_2^+$ ) and 146 amu ( $\text{C}_5\text{H}_8^+(\text{C}_6\text{H}_6)$ ) for integration times of 0.33 s, 0.33 s, and 1.65 s, respectively. For the conditions of this study, the ratio of the benzene dimer cation signal (100 kHz) to the benzene monomer cation signal (50 kHz) was  $\sim 2 : 1$ . The isoprene signal was normalized to the sum of the reagent ion signals

### Intercomparison of two isoprene measurement techniques

Leibrock et al.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

---

**Intercomparison of  
two isoprene  
measurement  
techniques**

---

Leibrock et al.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

to account for variations in the ion source intensity. The sensitivity of the instrument was calibrated by performing standard additions of a known amount of isoprene to the inlet flow. The calibration gas was a gravimetrically prepared 10 ppmv mixture of isoprene in synthetic air (see below). The background at the isoprene product mass was measured by periodically overflowing the inlet with dry, isoprene free synthetic air (Scott Specialty Gases). Mixing ratios of isoprene were obtained by subtracting the background levels from the normalized isoprene signal and then dividing by the measured calibration factor. The calibration factor was of the order of 2 Hz per pptv of isoprene at a total reagent ion signal of 150 kHz. The measured background signals were 100–200 Hz and the deviation on this signal was essentially determined by counting statistics. The detection limit ( $S/N = 3$ ) for these conditions was 21 pptv for a 1 s integration period.

### 2.3. GC

The GC was housed in a separate trailer on the mesa. Air samples for the GC were drawn through a 3 m long  $\times$  1/4" OD Teflon tube at approximately 12 SLPM. The end of the tube was also located at a height of  $\sim$  2 m above the ground within 5 m of the CIMS inlet. The in situ GC method is described in detail elsewhere (Goldan et al., 2000) and is only briefly outlined here. The GC utilized a 70 SCCM (standard cubic centimeters per minute) fraction of the whole air sample stream, passed it through an Ascarite trap at ambient temperature that removed carbon dioxide and ozone and subsequently through a cold zone held at  $-50^{\circ}\text{C}$  to remove most of the water. Hydrocarbons in the sample stream were subsequently concentrated for a 5 minute period in a cold trap that consisted of a 0.53 mm ID  $\times$  20 cm long section of  $\text{Al}_2\text{O}_3$  capillary column thermostatted at  $-170^{\circ}\text{C}$ . The concentrated sample was then injected onto a 0.53 mm ID  $\times$  50 m long analytical column of the same material (Chrompack Inc.) by heating the trap to  $100^{\circ}\text{C}$  in  $\sim$  6 s. Analysis was performed with a hydrogen carrier gas at a flow rate of 4 SCCM while the column temperature was ramped from 75 to  $150^{\circ}\text{C}$  at a rate of  $13^{\circ}\text{C}$  per min. The detection limit for isoprene (defined as signal to noise ratio = 2) was 5 pptv. The instrument had a 15 min cycle time allowing the acquisition

---

**Intercomparison of  
two isoprene  
measurement  
techniques**Leibrock et al.

---

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

and analysis of 4 samples per h. Between samples, the water trap was heated to 50°C and both it and the Ascarite trap were flushed with hydrocarbon free “zero air” for cleaning. The sensitivity of the FID was calibrated using gravimetrically prepared calibration standards of isoprene in synthetic air at the 10 ppmv level. For instrument calibration, these standards were diluted to the 0.1 to 2 ppbv range in a synthetic air matrix, to which ambient levels of carbon dioxide and water vapor were added, by a dynamic dilution system. The same gravimetrically prepared calibration standards were used by the CIMS instrument.

#### 2.4. Ancillary measurements

Ancillary measurements of chemical species included NO (chemiluminescence), NO<sub>2</sub> (photolytic conversion / chemiluminescence), total reactive nitrogen NO<sub>y</sub> (gold catalytic converter / chemiluminescence), and ozone (UV absorption). Wind speed and direction were also measured at the field site from a 10 m tower. The computer clocks for all the measurements were synchronized every morning and were found to deviate less than 5 s over the course of a day. Temperature, barometric pressure, and relative humidity were obtained from the Mesa Lab Weather Station (1855 m asl) at the National Center for Atmospheric Research located approximately 5 km south of Enchanted Mesa (NCAR, 1998).

### 3. Results

The CIMS and GC data were compared by averaging the individual 1.65 s CIMS data points over the corresponding 5 min GC sampling period. Figure 2 is a time series plot of the measurements obtained with both instruments as well as the time averaged CIMS data. The error bars on the averaged CIMS data represent the variability (standard deviation) of the CIMS isoprene measurements during the GC sampling period. Figure 3 shows the correlation of all averaged CIMS data with the individual GC mea-

---

**Intercomparison of  
two isoprene  
measurement  
techniques**

---

Leibrock et al.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[⏪](#)[⏩](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

surements. While there is a good overall correlation ( $R = 0.78$ ), the CIMS isoprene measurements clearly were consistently higher than those made with the GC. The linear regression of CIMS vs. GC data has a positive intercept of 67 pptv and a slope of 1.57. These data strongly indicate that the CIMS method suffered from an interference that led to over-measurement of isoprene in addition to a potential background problem. In an attempt to characterize the interference, the normalized differences (ND) between the CIMS and GC data were compared with both the chemical and meteorological ancillary measurements. The ND is a measure of the relative deviation between the measurements and is defined as:

$$ND = \frac{[CIMS] - [GC]}{\sqrt{[CIMS][GC]}} \quad (2)$$

where [CIMS] and [GC] are the concurrent isoprene measurements made by the CIMS and GC, respectively.

Of the meteorological parameters considered, the normalized differences were found to have a significant correlation only with wind direction (Figs. 4 and 5). Clearly, the ND was largest when the wind was from the south (Fig. 5). This indicates that an interference to the CIMS measurement was present in air masses transported from the Denver metropolitan area. The CIMS isoprene measurements are also remarkably well correlated with  $NO_x$  measurements ( $R = 0.7$ , see Fig. 6) even though  $NO_x$  and isoprene have very different sources. This suggests that the interfering compound to the CIMS measurements and nitrogen oxides were transported to the site from a collocated, anthropogenically influenced source.

#### 4. Conclusions

Unfortunately, the data obtained in this study demonstrate that the CIMS technique in its present form is not a viable method for isoprene measurements in or near an urban environment unless there are high levels of isoprene present. This is likely due

---

**Intercomparison of  
two isoprene  
measurement  
techniques**Leibrock et al.

---

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

to unidentified components in urban air that also react with the benzene reagent ions to give the same product mass as isoprene. We have demonstrated that 2-methyl-3-butene-2-ol (MBO) and structural isomers of isoprene such as 1,3-pentadiene also react with  $C_6H_6^+$  and  $(C_6H_6)_2^+$  to give mass 146 (Leibrock and Huey, 2000). However, we could not yet determine if these were the interfering species in this work. This study also demonstrates the largest difficulty with measurements of VOCs by positive ion chemical ionization: the ambient mix of organic compounds in polluted air is so complex that it is difficult for ion chemistry to selectively detect a tractable number of compounds.

The conditions of this intercomparison were very challenging for the CIMS method because of the low isoprene mixing ratios (less than 1 ppbv for most of the time) and the complex mix of VOC at the site due to local urban sources. These considerations suggest that the present CIMS isoprene method may be a viable technique for measurements in remote areas with high isoprene emissions and little anthropogenic influence. Sensitivity and time response of the technique were also clearly demonstrated in this study (Fig. 2); they indicate that the technique is a possibility for flux measurements of isoprene in remote areas using the eddy correlation technique (Lenschow et al., 1994).

One improvement to enhance the selectivity of the current technique would be the replacement of the quadrupole mass filter with a system that is capable of tandem mass spectrometry such as an ion trap mass spectrometer, (e.g. Dawson, 1976; March and Hughes, 1989). This would add another level of selectivity and allow to differentiate between isoprene and compounds that form products with equal masses, but have different internal structures.

*Acknowledgements.* E. L. was supported by a research grant by the German Academic Exchange Service (DAAD).



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## Intercomparison of two isoprene measurement techniques

Leibrock et al.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Print Version

Interactive Discussion

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## Intercomparison of two isoprene measurement techniques

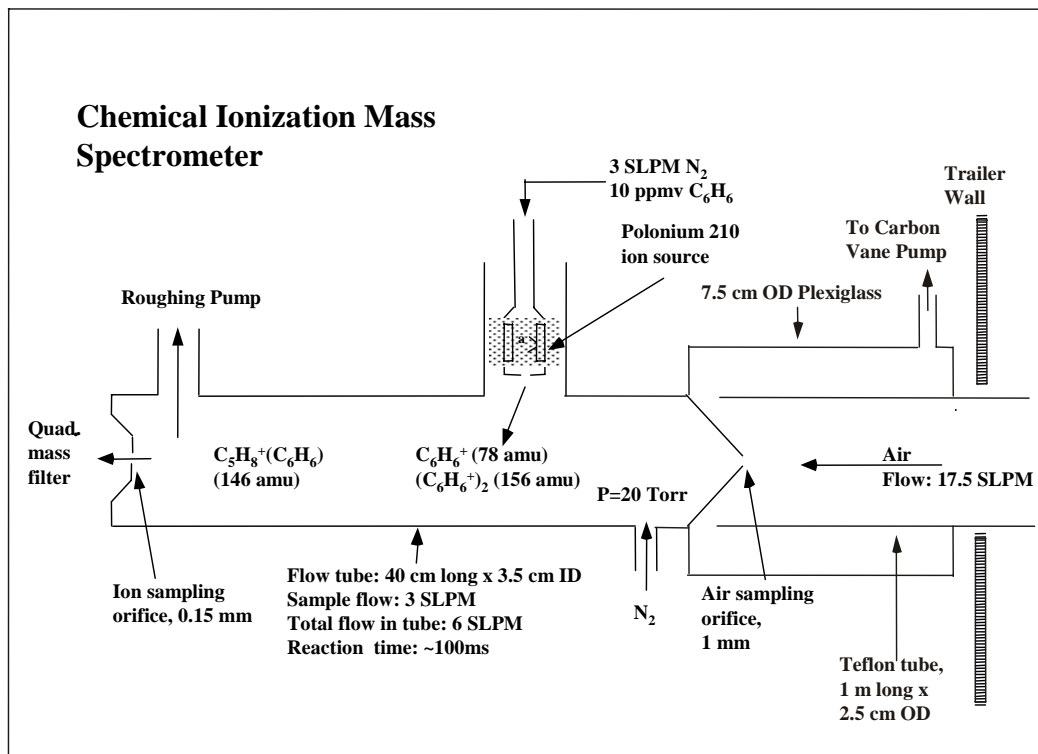
Leibrock et al.

---

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[⏪](#)[⏩](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

## Intercomparison of two isoprene measurement techniques

Leibrock et al.



**Fig. 1.** Simplified schematic drawing of the chemical ionization mass spectrometer (CIMS)

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

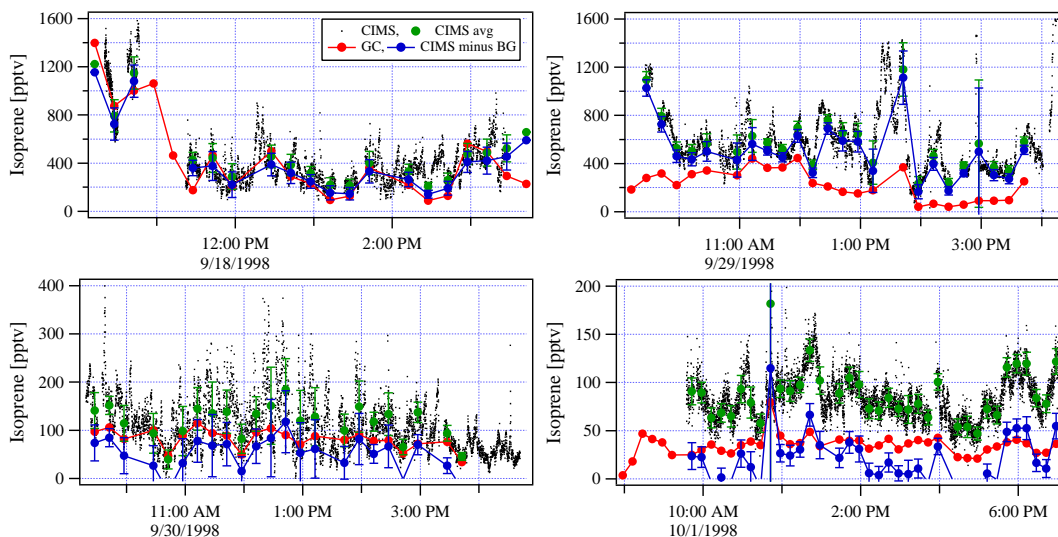
Full Screen / Esc

Print Version

Interactive Discussion

Intercomparison of  
two isoprene  
measurement  
techniques

Leibrock et al.



**Fig. 2.** Diurnal trends of isoprene measured with both instruments (CIMS: green dots, GC: red markers) during the intercomparison period as well as the CIMS data averaged (green markers) to the GC sampling periods. The CIMS background of about 70 pptv determined from Fig. 3 was subtracted from the CIMS averages resulting in the CIMS net signal (blue markers). The error bars on the CIMS data represent the variability (standard deviation) of the CIMS data during the GC sampling interval. Time is Mountain Daylight Time (MDT).

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

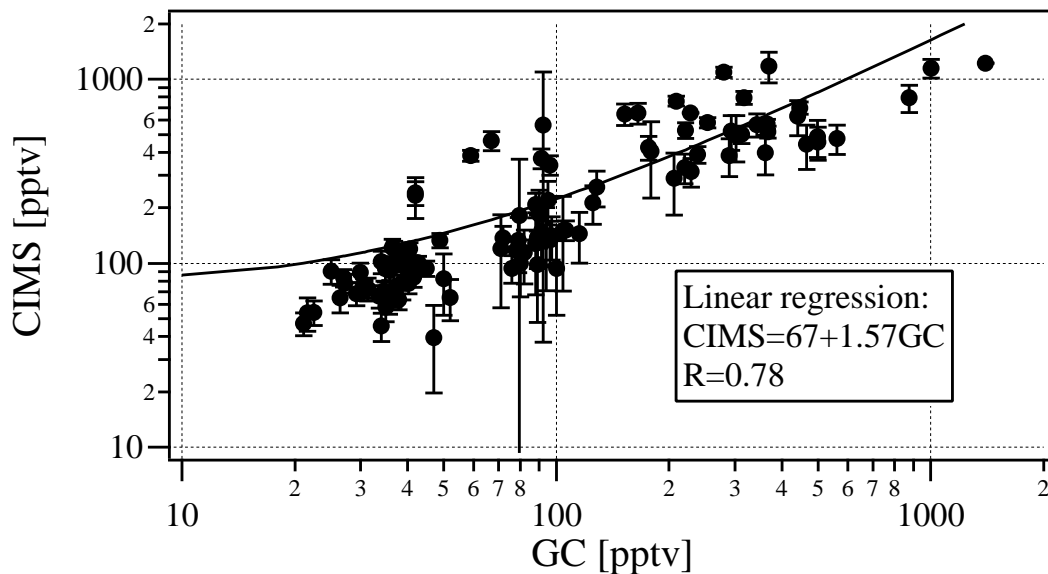
Full Screen / Esc

Print Version

Interactive Discussion

Intercomparison of  
two isoprene  
measurement  
techniques

Leibrock et al.

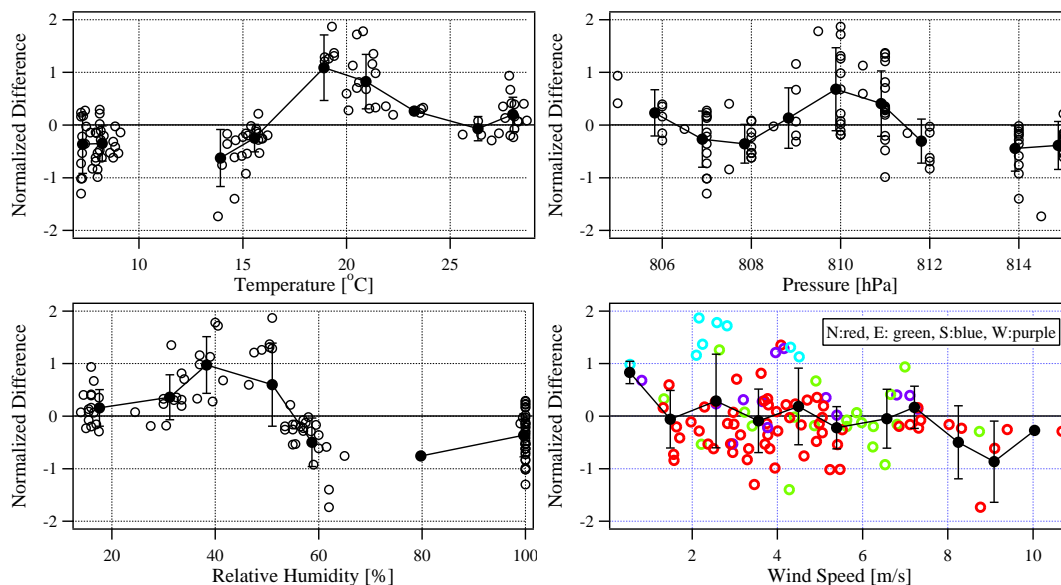


**Fig. 3.** Correlation of all GC measurements made during the intercomparison period with the averaged CIMS data. The error bars represent the standard deviation of the CIMS data during the GC sampling intervals. The solid line is the linear regression of CIMS vs. GC data, equation and correlation coefficient  $R$  are given in the text box: The intercept of about 70 pptv indicates an enhanced background in the CIMS measurements.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

## Intercomparison of two isoprene measurement techniques

Leibrock et al.



**Fig. 4.** Normalized differences of CIMS and GC data versus various meteorological parameters: temperature, pressure, relative humidity, wind speed, and wind direction. The CIMS data were averaged over the GC sampling time of 5 min and background corrected (open circles). To better visualize possible trends, the data were grouped into separate bins with a suitably chosen width and averaged within each bin (filled circles). The bars show the data variability (standard deviation) within each bin. Temperature, pressure, and relative humidity data were obtained from [NCAR \(1998\)](#).

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

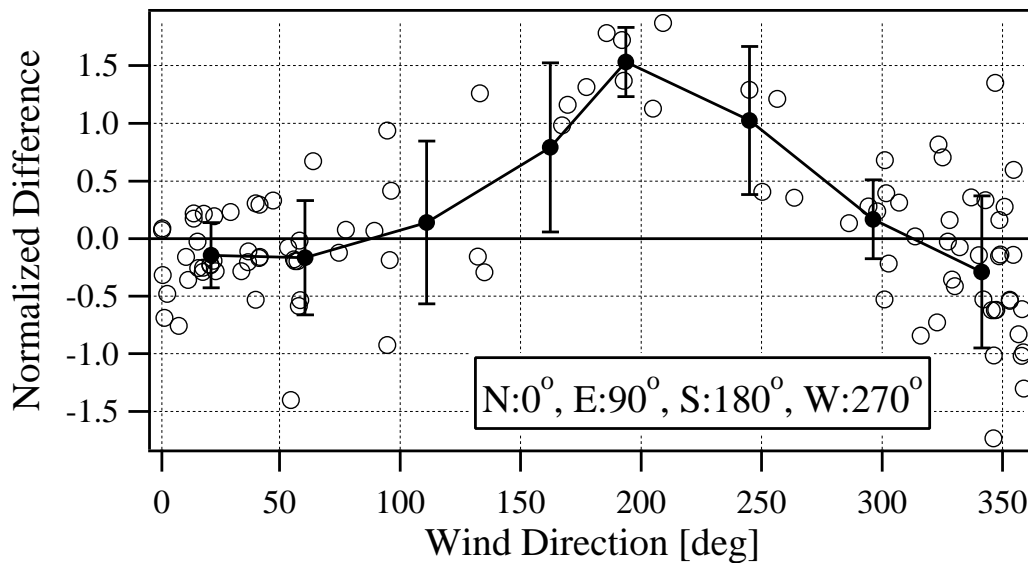
Full Screen / Esc

Print Version

Interactive Discussion

**Intercomparison of  
two isoprene  
measurement  
techniques**

Leibrock et al.

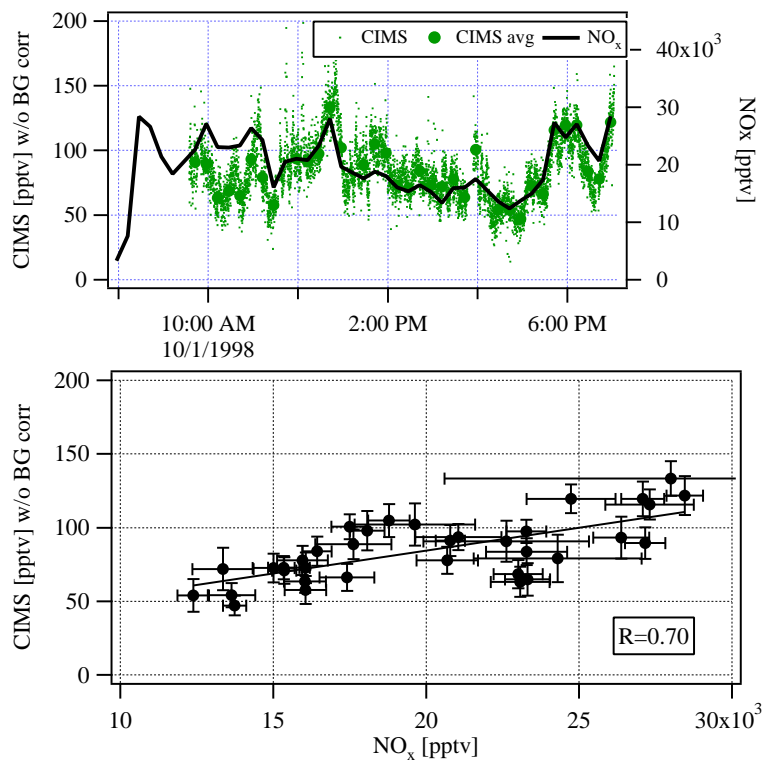


**Fig. 5.** Normalized differences of CIMS and GC data versus wind direction ( $0^\circ=N$ ,  $90^\circ=E$ ,  $180^\circ=S$ ,  $270^\circ=W$ )

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)

Intercomparison of  
two isoprene  
measurement  
techniques

Leibrock et al.



**Fig. 6.** The upper panel shows diurnal trends (1 Oct 1998) of non-background corrected isoprene data measured with the CIMS and their 5-min averages (left axis) and NO<sub>x</sub> (right axis) based on measurements of NO and NO<sub>2</sub> with a time resolution of 1 s each, suggesting a correlation between the data shown. The lower panel shows the linear correlation of the averaged, non-background corrected CIMS data versus the corresponding 5-min averages of the NO<sub>x</sub> values.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Print Version](#)[Interactive Discussion](#)