

1 **Second Response to Anonymous Referee #1**

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3 We thank the referee for another careful reading of our manuscript and the constructive comments that
4 will improve the final publication. We have addressed each comment with modifications in the
5 manuscript, and our responses to the reviewer are itemized below in green text following each specific
6 comment. A major revision to the paper is a new paragraph at the beginning of section 3.4 explaining
7 why we are presenting only shortwave radiative forcing values.
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Suggestions for revision or reasons for rejection (will be published if the paper is accepted for final publication)

The authors have provided (almost) satisfactory answers to my comments and requests for corrections. The manuscript could be further improved by being more precise and upfront in a few places:

- page 19, line 14: "total heating" is a little unclear. I think the authors mean "total heating by gases and aerosols". It would be clearer that way.

We have followed this suggestion and modified the text to read "total heating by gases and aerosols".

- Caption or text in Table 2 should repeat that heating rates are for SW and LW radiation combined.

We modified the last line of Table 2 to read "LW+SW Heating Rate (K day^{-1}) in tropical lower stratosphere"

- page 1, line 18: according to the usual convention, the SW RF should be negative (that is -1.2 W m^{-2}).

We have replaced " 1.2 W m^{-2} " with " -1.2 W m^{-2} " in the abstract.

- I couldn't find a justification in the text as to why the LW radiative forcing is ignored. It is clear from the abstract, page 16 line 18, and Table 2 that the numbers provided are for SW only. But the authors should be more upfront about this. I don't find the reason provided in the authors' response very convincing. The LW RF can be important to consider (eg for sulphate stratospheric aerosols) but is ignored here for simplicity (and because it is possibly less for some of the new material considered here).

We agree with the referee that a more explicit statement regarding our focus on shortwave radiative forcing would improve the presentation. We have added a new paragraph to clarify our approach in this paper at the beginning of Section 3.4, "Radiative forcing assessment", and a new sentence on page 17. Additionally, we note the lack of a clear convention on radiative forcing calculations for stratospheric aerosol geoengineering in the existing literature. For example, Ferraro et al. (2011) compute radiative forcing as the sum of longwave and shortwave components. In contrast, Pope et al. (2012) compute the increase in planetary shortwave albedo, which is equivalent to the shortwave radiative forcing divided by the incident solar flux. Though our tables and figures present only SW RF, we do discuss the approximate adjustment that the LW component would contribute to the RF (~10% for alumina, ~20% for sulfate, insignificant for diamond).

The new text at the beginning of section 3.4 reads as follows:

"In this work, we confine our assessment of radiative forcing to reflected solar radiation, acknowledging that absorption of thermal infrared radiation generates a longwave radiative forcing as well. For some

materials, this longwave forcing can be significant, reaching approximately 20% of the shortwave radiative forcing in the case of sulfate. Compared to the shortwave forcing, however, the longwave forcing is a much more sensitive function of profiles of atmospheric temperature, humidity, ozone, and well-mixed greenhouse gases. Furthermore, the instantaneous longwave forcing adjusts in response to changes in the temperature profile caused by radiative heating rate perturbations due to the aerosol particles. We therefore restrict our analysis to shortwave radiative forcing to narrow the scope of radiative transfer calculations and to reduce the number of radiative forcing values presented.”

And later in that section, at the end of page 17, we include:

“As stated at the beginning of this section, our analysis focuses on shortwave radiative forcing, as it provides for a less ambiguous comparison, relative to shortwave plus longwave forcing, of the efficacy of different particles in offsetting surface warming. However, we note that the longwave radiative forcing is about 10% of the shortwave RF for alumina, though of opposite sign, and is negligible for diamond.”

Even though we are able to state that the LW RF from alumina and diamond are small, comparisons with sulfate RF which ignore the LW are somewhat misleading. We have added a footnote to Table 2 stating that: “Sulfate fluxes needed to achieve a given RF would increase by approximately 20% if the LW RF component were included, decreasing the RF efficiencies shown here.”

A manuscript is in preparation by John Dykema that addresses both the shortwave and longwave radiative forcings and analyzes the stratospheric temperature response resulting from radiative heating rate perturbations caused by a variety of solid aerosol types.

- I still find the y-label to figure 6a confusing. The caption makes it clear but what is plotted is NOT a cross section. It would be more appropriate to call this quantity an "upscatter factor" in reference to the more usual quantities of extinction or scattering factors.

We have replaced the y-axis label in Figure 6a with “upscatter efficiency”, which we define in the caption as upscatter cross-section divided by geometric cross-section.

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1 Solar geoengineering using solid aerosol in the stratosphere

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7

8 Abstract

9 Solid aerosol particles have long been proposed as an alternative to sulfate aerosols for solar
10 geoengineering. Any solid aerosol introduced into the stratosphere would be subject to
11 coagulation with itself, producing fractal aggregates, and with the natural sulfate aerosol,
12 producing liquid-coated solids. Solid aerosols that are coated with sulfate and/or have formed
13 aggregates may have very different scattering properties and chemical behavior than do uncoated
14 non-aggregated monomers. We use a two-dimensional chemistry-transport-aerosol model to
15 capture the dynamics of interacting solid and liquid aerosols in the stratosphere. As an example,
16 we apply the model to the possible use of alumina and diamond particles for solar
17 geoengineering. For 240 nm radius alumina particles, for example, an injection rate of 4 Tg yr⁻¹
18 produces a global-average shortwave radiative forcing of -1.2 W m⁻² and minimal self-
19 coagulation of alumina although almost all alumina outside the tropics is coated with sulfate. For
20 the same radiative forcing, these solid aerosols can produce less ozone loss, less stratospheric
21 heating, and less forward scattering than do sulfate aerosols. Our results suggest that
22 appropriately sized alumina, diamond or similar high-index particles may have less severe
23 technology-specific risks than do sulfate aerosols. These results, particularly the ozone response,
24 are subject to large uncertainties due the limited data on the rate constants of reactions on the dry
25 surfaces.

26

27 1 Introduction

1 Solar geoengineering, or Solar Radiation Management (SRM) is the possibility of deliberately
2 introducing changes to the Earth’s radiative balance to partially offset the radiative forcing of
3 accumulating greenhouse gases and so lessen the risks of climate change. Most research on SRM
4 has concentrated on the possibility of adding aerosols to the stratosphere, and essentially all
5 atmospheric modeling of stratospheric aerosol injection has focused on increasing the loading of
6 aqueous sulfuric acid aerosols (Rasch et al. 2008; Heckendorn et al., 2009; Niemeier et al., 2011;
7 Pitari et al., 2014). The possibility that solid aerosol particles might offer advantages over
8 sulfates, such as improved scattering properties, was first suggested almost two decades ago, but
9 analysis has been almost exclusively limited to conceptual studies or simple radiative transfer
10 models (Teller et al., 1997; Blackstock et al., 2009; Keith, 2010; Ferraro et al., 2011; Pope et al.,
11 2012).

12 Any solid aerosol injected directly into the stratosphere for geoengineering purposes would be
13 subject to coagulation with itself and with the natural background or volcanic sulfate aerosol.
14 Aggregates of solid aerosols have very different physical structure and scattering properties than
15 do liquid sulfate aerosol particles. The lifetime and scattering properties of a solid aerosol are
16 strongly dependent on these dynamical interactions, and the chemical properties of the aerosol
17 depend on the extent to which it becomes coated by the ambient sulfate.

18 We have modified the Atmospheric and Environmental Research (AER) two-dimensional
19 chemistry-transport-aerosol model (Weisenstein et al., 2004; 2007) to capture the dynamics of
20 interacting solid and liquid aerosols in the stratosphere. Our model now includes a prognostic
21 size distribution for three categories of aerosols: liquid aerosols, solid aerosols, and liquid-coated
22 solid aerosols. The model’s coalescence kernel has been modified and extended to parameterize
23 the interactions of particles across size bins and between all combinations of the three categories.
24 The surface area, sedimentation speed, and coalescence cross-section of an aggregate of solid
25 particles depend on the geometry of the aggregate. The model parameterizes this physics using a
26 fractal dimension and allows that fractal dimension to change with age or with a liquid coating.
27 The chemistry and aerosol schemes in the model are interactive, while dynamical fields are
28 prescribed.

29 Turning now to the context of this work, it is useful to divide overall consideration of the risks
30 and efficacy of SRM into two components. First, the ability, or *efficacy*, of idealized SRM—

1 conceived as a reduction in the solar constant—to compensate for the risks of accumulating
2 greenhouse gases. And, second, the *technology-specific risks* of any specific engineered
3 intervention that produces a change in radiative forcing. Uncertainty in the efficacy of SRM, the
4 first component, rests on uncertainty in the climate’s large-scale response to forcing. Results
5 from a large set of climate models suggest that idealized SRM can do a surprisingly good job in
6 reducing model-simulated climate changes, both locally and globally, which, in our view, is a
7 primary motivation for continued research on SRM (Kravitz et al., 2014; Moreno-Cruz et al.,
8 2011).

9 Evaluation of the technology-specific risks depends on the specific technology. For sulfate
10 aerosols these risks include, but are not limited to, (a) ozone loss, (b) radiative heating of the
11 lower stratosphere which causes changes in atmospheric temperature and dynamical transport,
12 and (c) the fact that sulfates produce a relatively high ratio of downward scattering to upward
13 scattering so that they substantially increase the ratio of diffuse to direct radiation (Kravitz et al.,
14 2012) which in turn may alter atmospheric chemistry and ecosystem functioning (Mercado et al.,
15 2009; Wilton et al., 2011). In addition to the risks, it may be difficult to produce sufficiently
16 large radiative forcings using SO₂ because of the decreasing efficiency at higher SO₂ inputs
17 (Heckendorn et al., 2009; English et al., 2012).

18 The use of solid particles for SRM offers the *potential* to address all of the limitations of sulfate
19 particles. Solid aerosols do not, for example, directly increase the stratospheric volume of the
20 aqueous sulfuric acid that drives hydrolysis reactions, an important pathway through which
21 sulfate aerosols cause ozone loss. In addition, some solid aerosols (e.g., diamond, alumina, or
22 titania) have optical properties that may produce less heating in the lower stratosphere (Ferraro et
23 al., 2011), and any solid with a high index of refraction can reduce forward scattering.

24 The use of solid aerosols, however, introduces new risks that require evaluation. The dry
25 surfaces of the solid aerosols, for example, may catalyze reactions that cause ozone loss (Tang et
26 al., 2014a; 2014b). This risk is hard to evaluate because the rates of many potentially important
27 chemical reactions remain unmeasured for substances such as diamond that are novel in the
28 stratosphere. Moreover, by spreading the natural background sulfuric acid over a larger surface
29 area as will occur when background sulfate coats the solid particles, the addition of solid aerosols
30 will increase reactions that depend on sulfate surface area density rather than sulfate volume.

1 Our motivation for studying solid particles is the possibility that they enable a decrease in the
2 risks of SRM (e.g., ozone loss) or an increase in its efficacy such as the ability to produce larger
3 radiative forcings, or an improved ability to “tune” the spectral or spatial characteristics of the
4 radiative forcing (Blackstock et al., 2009; Keith, 2010). This is in contrast to much of the prior
5 literature that has focused on the potential of solid particles to deliver higher mass-specific
6 scattering efficiency, thus reducing the amount of material needed to produce a given radiative
7 forcing. We do not see this as an important motivation as it appears that the cost of lofting
8 materials to the stratosphere is sufficiently low that cost is not an important barrier to
9 implementation of SRM (McClellan et al., 2012).

10 In this paper, our focus is on developing the tools and methodology for assessing the risks and
11 performance of solid particles injected into the stratosphere for SRM. The tool described here is
12 a new solid-liquid stratospheric aerosol model, and the methodology is a comparison of
13 environmental side-effects such as ozone loss and forward scattering as a function of the global
14 radiative forcing. We use aluminum oxide (alumina) aerosol as the primary example. Diamond
15 appears to be superior to alumina in several respects, perhaps the most important being that it has
16 minimal absorption in the thermal infrared. We examine diamond, but choose alumina as the
17 primary example because there is a broad basis to examine alumina’s potential environmental
18 impacts. Unlike many other solid particles proposed for SRM, there is prior work examining
19 alumina’s impacts on stratospheric chemistry (Danilin et al., 2001; Jackman et al., 1998; Ross
20 and Shaeffer, 2014), work that was produced from NASA-funded studies starting in the late
21 1970’s motivated by concerns about the ozone impact of space shuttle launches (alumina is a
22 major component of the shuttle’s solid rocket exhaust plume). Moreover, alumina is a common
23 industrial material with a high index of refraction for which there is substantial industrial
24 experience with the production of nanoparticles (Hinklin et al., 2004; Tsuzuki and McCormick,
25 2004). With respect to potential environmental impacts of alumina deposition on Earth’s surface,
26 the fact that aluminum oxides are a common component of natural mineral dust deposition
27 provides a basis for assessing impacts (Lawrence and Neff, 2009). For diamond, there is
28 evidence that diamond nanoparticles are nontoxic to biological systems (Shrand et al., 2007). A
29 much more substantive assessment of the human health and ecosystems impacts of any proposed
30 solid aerosol would be required, however, prior to serious consideration of their use for
31 geoengineering.

1 The remainder of this paper is organized as follows. The solid-liquid model is presented in
2 Section 2, results for geoengineering injection of alumina and diamond in Section 3, and
3 discussion in Section 4.

4

5 **2 Aerosol model**

6 We have incorporated solid aerosols into the AER 2-D chemistry-transport-aerosol model
7 (Weisenstein et al., 1997; 2004; 2007). The aerosol module, which employs a sectional scheme,
8 has been modified to include three separate classes of aerosols, each with its own size
9 distribution: solid particles, liquid H₂SO₄-H₂O particles, and mixed solid-liquid particles. To
10 fully specify the mixed particles we keep track of the volume of liquid H₂SO₄-H₂O solution
11 coating the mixed particles. Unlike liquid particles that coagulate into larger spheres, solid
12 particles coagulate into fractal structures with more complex properties. The fractal properties
13 are required to predict the effective size of the particles appropriate to determining coagulation
14 interactions and gravitational settling. Fractal properties are also needed to determine the
15 condensation rate of H₂SO₄ gas onto alumina particles and the aerosol surface area density that is
16 important to heterogeneous chemistry and ozone depletion.

17 The AER 2-D model includes standard chemistry relevant to ozone (Weisenstein et al., 2004) as
18 well as aerosol microphysics and the relevant sulfur chemistry (Weisenstein et al., 1997; 2007).
19 The model includes sulfur-bearing source gases dimethyl sulfide (DMS), CS₂, H₂S, OCS, and
20 SO₂ emitted by industrial and biogenic processes as well as the product gases methyl sulfonic
21 acid (MSA), SO₂, SO₃, and H₂SO₄. Chemical reactions affecting sulfur species are listed in
22 Weisenstein et al. (1997) and their rates have been updated according to Sander et al. (2011).
23 Values of OH and other oxidants are calculated interactively along with ozone and aerosols
24 (Rinsland et al., 2003; Weisenstein et al., 2004). Further description of the chemistry directly
25 relevant to ozone is included in Section 3.5. The model's 2-D transport is prescribed based on
26 calculations by Fleming et al. (1999) for each year from 1978 to 2004, employing observed
27 temperature, ozone, water vapor, zonal wind, and planetary waves. Different phases of the
28 quasi-biennial oscillation (QBO) are included in the observational data employed. We average
29 the transport fields over the years 1978-2004 into a climatology and employ that circulation each
30 year of our ten year calculations. Temperature fields are also prescribed based on climatological

1 observations for the same averaging period. The domain is global, from the surface to 60 km,
2 with resolution of 1.2 km in the vertical and 9.5 degrees in latitude. Though the model is
3 primarily suited to modeling the stratosphere and upper troposphere, it does contain a
4 parameterization of tropospheric convection (Dvortsov et al., 1998) that serves to elevate SO₂
5 concentrations in the tropical upper troposphere.

6 The AER 2-D aerosol model, along with several other 2-D and 3-D models, was evaluated and
7 compared to observations in SPARC (2006). The AER model was found to reasonable represent
8 stratospheric aerosol observations in both nonvolcanic conditions and in the period following the
9 eruption of Mt Pinatubo. Noted deficiencies, common to most models, included too high
10 calculated values of aerosol extinction in the tropics between the tropopause and 25 km as
11 compared to SAGE II extinctions at 0.525 and 1.02 μm during nonvolcanic periods. The growth
12 and decay of the stratospheric aerosol layer following the Mt. Pinatubo eruptions was generally
13 well-represented by the AER model as compared to lidar and satellite observations from 1991 to
14 1997, though uncertainties in the initial SO₂ injection amount and vertical distribution limit our
15 interpretation. Dynamical variability on short time scales was underestimated by the model.

16 Sulfate aerosol formation is thought to be initiated mainly by binary homogeneous nucleation of
17 H₂SO₄ and H₂O vapors, primarily in the tropical tropopause region. The aerosol size distribution
18 is modified by condensation and evaporation of gas-phase H₂SO₄ and by coagulation among
19 particles (Brock et al. 1995; Hamill et al., 1997). Sulfate aerosol particles are assumed to be
20 liquid spheres with equilibrium composition (H₂SO₄ and H₂O fractions) determined by the local
21 grid box temperature and water vapor concentration (Tabazadeh et al., 1997). The model uses a
22 sectional representation of particle sizes, with 40 logarithmically-spaced sulfate aerosol bins,
23 representing sizes from 0.39 nm to 3.2 μm, with aerosol volume doubled between adjacent bins.
24 Particle distributions are also modified by sedimentation and by rainout/washout processes in the
25 troposphere. The sedimentation formulation is described below. Rainout/washout process are
26 represent by a first order loss term in the troposphere with removal lifetime ranging from 5 days
27 at the surface to 30 days at the tropopause.

28 Solid particles are modeled with a similar sectional representation; in this case it is the number of
29 monomers per particle that is doubled in successive bins. Only the monomers, the primary
30 particles directly injected into the atmosphere, are assumed spherical. Larger particles produced

1 by coagulation assume fractal structures that obey a statistical scaling law where the fractal
2 dimension D_f determines how the size of an aggregate of particles is related to the number of
3 primary particles (Filippov et al., 2000; Maricq, 2007). . The radius of gyration R_g of a fractal
4 (the root-mean-square distance from the center of mass) is given by:

$$5 \quad R_g = R_0 (N_i/k_f)^{(1/D_f)}$$

6 where R_0 is the primary particle radius, N_i the number of monomers in the fractal of bin i , and k_f
7 is a prefactor. Thus particle mass is proportional to $R_g^{D_f}$. The fractal dimension D_f for a given
8 material has been found to be invariant for a wide range of R_0 and N_i values.

9 The surface area (S) for a fractal particle can be parameterized with an effective radius R_{eff} which
10 can be related to primary radius and the number of monomer cores in the particle:

$$11 \quad R_{\text{eff}} = R_0 (N_i/k_h)^{(1/D_h)}$$

$$12 \quad S_i = (4 \pi R_0^2) * (N_i/k_h)^{(2/D_h)}$$

13 where D_h and k_h are the scaling exponent and prefactor specific to surface transfer processes.

14 With fractal dimension $D_f < 2.0$, D_h can be assumed equal to 2.0. With $D_f > 2.0$, D_h can be
15 assumed equal to D_f (Filippov et al., 2000). When $D_h = 2.0$, the surface area of a fractal particle
16 is equal to the surface area of the monomer multiplied by the number of monomers in the
17 aggregate. This formalism is most appropriate for large values of N_i (i.e., greater than 100). For
18 consistency at small values of N_i , we assume that $k_f = k_h = 1$, since we find that simulations
19 producing only small N_i values are most efficient for geoengineering.

20 The solid particles are allowed to interact with background stratospheric sulfate particles by
21 coagulation, and with gas phase H_2SO_4 and H_2O by condensation and evaporation. We use R_g as
22 the particle radius when calculating the coagulation kernel, the probability that two particles will
23 combine into one on collision (Maricq, 2007). The coagulation formulation between and among
24 different particle types is detailed in Appendix A. The condensation rate, also detailed in
25 Appendix A, depends on particle surface area, and secondarily, on a radius of curvature for the
26 Kelvin correction. We use R_0 as the radius-of-curvature in the condensation equation, since gas
27 molecules see the individual monomers making up the fractal. We model mixed-phase particles
28 by tracking particle number per bin and mass of H_2SO_4 per bin in the mixed particles. Volume
29 and surface area of the mixed particles depends also on the H_2O present in the equilibrium

1 H₂SO₄-H₂O solution. Above about 35 km, coated particles will lose their sulfate coating by
2 evaporation and become dry again.
3 The sedimentation velocity of fractal particles represents a balance between the gravitational
4 force, proportional to particle mass, M_p, and the drag force, proportional to the particle velocity
5 and the two-dimensional surface area projection of the particle, A_{2D}, and inversely proportional
6 to the particle radius R_p. Sedimentation velocity is modified by the Cunningham slip-flow
7 correction, G, which accounts for larger sedimentation velocities with lower air density (Seinfeld
8 and Pandis, 2006). We obtain sedimentation velocity W_{sed} from

$$9 \quad W_{\text{sed}} = (M_p g R_p G) / (6 \eta A_{2D}),$$

10 where η is the viscosity of air and g the gravitational constant. For spheres, W_{sed} is proportional
11 to R_p². For all fractal cases, R_p is taken to be R_g, and with D_f ≥ 2, the area projection is taken to
12 be πR_g^2 yielding a W_{sed} proportional to G*N_i^{(D_f-1)/D_f}. With D_f < 2, the fractal is porous and the
13 area projection is N_i*πR₀², yielding a W_{sed} proportional to G*N_i^(1/D_f) (Johnson et al., 1996). For
14 coated particles, the particle mass, M_p, is the sum of the solid particle mass and H₂SO₄-H₂O
15 mass, and particle radius is taken to be R_g increased by the thickness of a uniform coating.
16 However, when the radius of a sphere enclosing the total particle volume is larger than R_g plus a
17 monolayer of H₂SO₄, we use the spherical radius rather than R_g.

18

19 **3 Model Results**

20 Before turning to the results, we use the following sub-section to describe (and provide some
21 rationale for) the solid aerosol particles that we choose as test cases, and then in Section 3.2 we
22 describe a few results regarding the sedimentation of aggregates that are useful in understanding
23 the model results.

24

25 **3.1 Test cases: Alumina and diamond aerosol particles**

26 Several prior studies have examined a range of possible solid aerosols and performed some
27 simple optimizations (Teller et al., 1997; Pope et al., 2012; Blackstock et al., 2009). For
28 simplicity we only considered spherical dielectric particles made of materials that have

1 negligible solubility in the aqueous sulfuric acid found under typical stratospheric conditions. An
2 ideal material for SRM would have (a) a high index of refraction, (b) a relatively low density, (c)
3 negligible absorption for both solar and thermal infrared spectral regions, and finally (d) it
4 should have well understood surface chemistry under stratospheric conditions. In addition, even
5 though this research is exploratory, materials are more plausible as candidates for deployment for
6 SRM if they have low and well understood environmental toxicity and if there is a track record
7 of production of industrial quantities of the material in the appropriate half micron size regime.

8 We chose alumina, or aluminum oxide (Al_2O_3), as our primary test case because it has a
9 relatively high index of refraction ($n=1.77$ in the middle of the solar band) and because there is a
10 substantial literature on its chemistry (Molina et al., 1997; Sander et al., 2011) and stratospheric
11 chemical impact (Danilin et al., 2001; Jackman et al., 1998). However, alumina has infrared
12 absorption bands in the thermal infrared that will reduce its net radiative forcing and will cause
13 some heating of the lower stratosphere (Ross and Schaeffer, 2014).

14 We chose diamond as a secondary test case because of its near-ideal optical properties: it has a
15 very high index ($n=2.4$) and negligible absorption for both solar and thermal infrared spectral
16 regions. Despite this we did not choose diamond as the primary test case because there is
17 minimal data about the chemistry of relevant compounds on diamond surfaces under
18 stratospheric conditions, and also, while industrial synthetic submicron diamond is now available
19 at under 100 USD per kilogram, there is still far less industrial heritage on diamond production
20 to assess the challenges of scaling production technologies to hundreds of thousands of tons per
21 year..

22 Alumina is an important industrial material as a precursor for aluminum production and for a
23 variety of uses from sunscreen compounds applied to the skin to industrial catalysis. The global
24 production rate is approximately 100 Tg yr^{-1} (USGS, 2014). There is a very large body of
25 experience in making alumina nanoparticles. For example, liquid-feed flame spray pyrolysis is
26 used to make structured nanoparticles of alumina in quantities greater than 1 kt yr^{-1} (Hinklin et
27 al., 2004). As we will see, the optimal size for a spherical alumina particles used as a scatterer in
28 the stratosphere is of order 200 nm radius. Most of the industrial effort is focused on making
29 smaller particles for catalysis but there are examples of production of relatively monodisperse
30 particles with radii greater than 50 nm (Hinklin et al., 2004; Tsuzuki and McCormick, 2004).

1 For the purposes of this paper we will assume that it is possible to make roughly spherical
2 alumina particles with a size range between 50 and 400 nm radius. This is a working assumption
3 that seems plausible based on the very large technical literature (> 1,000 papers in the last decade
4 on alumina nanoparticles) and industrial base devoted to production of these materials. But it is
5 simply an assumption. A significant effort involving experts from industry and academia would
6 be required to meaningfully assess the difficulty of producing large quantities of alumina with a
7 suitable size and morphology for solar geoengineering.

8 There is rapidly growing industrial production of sub-micron diamond powders (Krueger, 2008),
9 so there is no doubt that particles with appropriate morphology can be produced. However, the
10 industrial production volumes and academic literature on production technologies are far smaller
11 than for alumina.

12 Table 1 lists the numerical experiments performed for this study with the AER 2-D model.
13 Simulations with alumina particles employ a range of injected monomer sizes, from 80 to 320
14 nm in radius. These simulations allow us to analyze the trade-offs between sedimentation rate,
15 radiative forcing, and ozone depletion. For diamond, we perform simulations only for injected
16 monomer sizes of 160 nm, near the radiative optimum. We also repeat simulations performed by
17 Pierce et al. (2010) for geoengineering injections of SO₂ and H₂SO₄. A range of injection rates is
18 used for each injected substance and each injected monomer radius to test linearity of the
19 response. Each scenario is calculated with a 10 year integration period, using dynamical fields
20 representing the 1978-2004 average repeated each year and fixed boundary conditions from
21 approximately the year 2000, until an annually-repeating result is achieved. We analyze results
22 from the final year of each calculation, concentrating on annual average conditions.

23

24 **3.2 Factors controlling settling of aggregates**

25 As discussed above, the dynamics of aggregated particles depend on their fractal dimension D_f .
26 No observational data on the fractal dimension of ~100 nm hard spheres aggregating under
27 stratospheric conditions is available. As a guide, we adopt the value of D_f obtained in studies of
28 the formation of fractal alumina aggregates from much smaller monomers at atmospheric
29 pressure produced by combustion and oxidation of liquid aluminum drops that can result from
30 burning solid rocket fuel. These studies, which produced aggregates of approximately 1 μm

1 composed of primary particles of a few tens of nanometers in diameter, determined that the
2 fractal dimension D_f for alumina is 1.60 ± 0.04 (Karasev et al., 2001; 2004), implying a sparsely-
3 packed fractal. For comparison, soot aggregates typically have D_f values of ~ 2.0 (Kajino and
4 Kondo, 2011; Maricq and Nu, 2004), while a value of 3.0 is appropriate for liquid particles
5 which remain spherical upon coagulation. The density of alumina particles is taken to be 3.8 g
6 cm^{-3} and that of diamond to be 3.5 g cm^{-3} . We assume the same fractal dimensions for diamond
7 as for alumina.

8 Sedimentation velocity strongly influences stratospheric lifetimes. Figure 1(a) shows annual
9 average sedimentation velocities in the tropics as a function of altitude for uncoated alumina
10 particles for monomer radii from 80 to 320 nm. Shown are sedimentation velocities for
11 individual monomers and for fractals with $N=4$ and $N=32$, all with fractal dimension $D_f=1.6$.
12 $N=32$ fractals are not shown for monomers larger than 160 nm because significant numbers of
13 such fractals do not form in our simulations, however we do show $N=256$ fractals with 80 nm
14 monomers. Alumina monomers fall at a faster rate than sulfate particles of the same diameter,
15 given their greater density (3.8 g cm^{-3} for Al_2O_3 , approximately 1.7 g cm^{-3} for stratospheric
16 $\text{H}_2\text{SO}_4\text{-H}_2\text{O}$ particles), and diamond particles (not shown) fall only slightly slower than alumina
17 particles of the same radius owing to 8% smaller density. Fractal particles fall faster than the
18 monomers they are composed of in the troposphere and lower stratosphere, but at the same rate
19 in the middle and upper stratosphere where the Knudsen number $\text{Kn} > 1$ and the slip-flow
20 correction has the opposite size dependence as the other terms. Figure 1 also shows the model's
21 average upward advective velocity in the tropics as a function of altitude for comparison. Where
22 sedimentation velocity exceeds average upwelling velocity, we expect alumina lifetime and
23 vertical extent to be greatly impacted. This occurs only above 30 km for 80 nm monomers, but
24 above 24 km for 160 nm monomers and 19 km for 240 nm monomers. For 240 nm and 320 nm
25 monomers injected into the tropics at 20-25 km altitude, only a fraction of the injected mass will
26 be lofted to higher altitudes and distributed to high latitudes by the Brewer-Dobson circulation.

27 It is known that soot particles, which form fractals similar to alumina particles, eventually
28 assume a more compact structure in the atmosphere after acquiring a liquid coating (Kajino and
29 Kondo, 2011; Mikhailov et al., 2006). Observations on the liquid uptake properties of alumina
30 and their potential shape compaction are not available. For simplicity, we assume that the
31 alumina particles are hydrophobic until they are coated with a sulfate-water mixture by

1 coagulation with existing sulfate particles, and then they may take up additional H_2SO_4 and H_2O
2 by condensation. The effects of this assumption are expected to be small under non-volcanic
3 conditions, as most (>95%) stratospheric sulfate mass exists in condensed form. To test the
4 potential effect of compaction of liquid-coated solid alumina particles, we perform additional
5 model calculations assuming that the wetted particles change their fractal dimension D_f from 1.6
6 to 2.8, and their surface area scaling exponent D_h from 2.0 to 2.8, likely the maximum
7 compaction that could be achieved. While a time lag from initial wetting to shape compaction
8 may be appropriate, we assume instantaneous compaction on wetting for calculations labeled
9 “compact coated” as a way to bracket the effect. When the compacted particles lose their H_2SO_4
10 by evaporation, they are assumed to retain their compact shape. Sedimentation velocities for
11 these coated and compacted particles are shown in Fig. 1(b). In this case, higher order fractals
12 fall at faster velocities than their respective monomers at all altitudes, which will affect the
13 residence time of alumina and its calculated atmospheric burden.

14

15 **3.3 Aerosol distribution and burden**

16 We model geoengineering by injection of alumina particles for a number of parametric model
17 scenarios to evaluate the effect of (1) injected particles size, (2) injection rate, and (3) the fractal
18 geometry of sulfate-coated alumina particles. For all scenarios, injection occurs in a broad band
19 from 30°S to 30°N and from 20-25 km in altitude. This is the same injection region used in
20 Pierce et al. (2010) and was chosen to maximize the global distribution and residence time of
21 geoengineered aerosols while minimizing localized injection overlaps. We assume that it is
22 feasible to emit alumina particle monomers with a uniform diameter, either by a flame process at
23 the injection nozzle or by releasing prefabricated particles. Particles are released continuously at
24 injection rates of 1, 2, 4, or 8 Tg per year, all as monomers of a single radius (80 nm, 160 nm,
25 240 nm, or 320 nm), as detailed in Table 1. Stratospheric particle injections are continuous in
26 time and the simulations are continued for 10 years until a steady atmospheric concentration is
27 reached. Alumina particles that become coated with sulfate are treated either as retaining their
28 sparse structure with fractal dimension D_f of 1.6 or instantaneously becoming more compact
29 fractal particles with D_f of 2.8. We use a 2-D model for computation efficiency in this first
30 evaluation of geoengineering by solid particle injection, and thus we implicitly mix the injected

1 material into zonally-uniform bands dictated by the model's spatial resolution of 9.5 degrees
2 latitude by 1.2 km altitude. The impact of this simplification, along with the neglect of enhanced
3 coagulation in injection plumes, will be discussed in Section 4.

4 We first examine the calculated concentration and size distribution of atmospheric alumina under
5 a geoengineering scenario with an injection rate of 1 Tg yr^{-1} , assuming no particle compaction on
6 coating with sulfate. Figure 2, top panels, shows the mass mixing ratio of alumina (ppbm) with
7 injections of 80 nm monomers and 240 nm monomers. Significant alumina concentration exist
8 up to 40 km altitude when 80 nm particles are injected, but only below 30 km for injection of
9 240 nm particles due to the difference in sedimentation speeds. The peak mass mixing ratio of
10 alumina with 80 nm monomers injected is 40% larger than that with 240 nm monomers injected.
11 The lower panels of Fig. 2 show the concentration of particles (cm^{-3}) for the same cases. Particle
12 concentrations of up to 25 cm^{-3} are found for 1 Tg yr^{-1} injection of 80 nm monomers but remain
13 less than 3 cm^{-3} for injection of 240 nm monomers. The particle concentration drops away from
14 the injection region as the monomers coagulate into fractals and have time to settle downward.
15 The low number densities with $R_0=240 \text{ nm}$ result in minimal coagulation between alumina
16 particles.

17 The distribution of stratospheric alumina mass into monomers and fractals is shown in Fig. 3 at
18 the equator with 80 nm monomers injected (panel a) and for the global average with 80, 160,
19 and 240 nm monomers injected (panels b-d), all with 1 Tg yr^{-1} of emissions. With injection of
20 80 nm monomers, 27% of the mass remains in monomers at the equator, with no more than 13%
21 of the mass in any size bin with 2 or more monomers in the fractal. Some fractal particles
22 comprised of 1024 monomers exist. At higher latitudes, the monomer fraction drops and the
23 proportion in higher order fractals increases, as seen in the global average (panel b). The fraction
24 of monomers coated, shown as the blue portion of each bar, increases with distance from the
25 tropical injection region. Coated fractions also increase with increasing numbers of monomers
26 per fractal particle. This reflects both the longer residence time of the larger particles and their
27 large cross-section, which enhances coagulation with sulfate particles. Virtually all of the
28 alumina mass is coated for fractals with more than 128 monomers per particle. Alumina in the
29 troposphere is almost all coated with sulfate due to the large sulfate concentrations there, though
30 alumina concentrations are small. With injection of 160 nm alumina monomers, 71% of the
31 global mass remains in monomers, and fractals composed of only 2-16 monomers are found in

1 significant concentrations. Results for diamond closely resemble those for alumina with 160 nm
2 monomers injected. With injection of 240 nm alumina monomers, 94% of the mass remains in
3 monomers, and with 320 nm monomers injected, 98% is monomers. Larger fractions of the
4 alumina mass are coated in these latter cases.

5 The mass fraction in monomers versus higher order fractals varies with injection rate. Figure
6 4(a) shows mass fraction vs the number of monomers per particle for injection of 80 nm
7 monomers at rates varying from 1 Tg yr⁻¹ to 8 Tg yr⁻¹. The 1 Tg yr⁻¹ cases (green lines) match
8 the global mass fractions shown in Figure 3. As the injection rate increases, the mass fraction of
9 monomers decreases while the peak distribution shifts to larger fractals. At injection rates of 2,
10 4, and 8 Tg yr⁻¹, the size distribution peaks at 32, 64, and then 128 monomers per particle, and
11 fractals composed of 2000 monomers are found. Figure 4(b) shows a similar figure with
12 injection of alumina as 160 nm radius monomers. Because these particles contain 8 times the
13 mass of the 80 nm monomers, particle concentrations are considerably smaller and coagulation
14 less effective. Fractals containing more than 128 monomers do not occur in significant
15 concentrations, even with 8 Tg yr⁻¹ of emission. For injection of 240 nm monomers (Fig. 4c),
16 70% of the particles remain as monomers even with 8 Tg yr⁻¹ of emission, and fractals exceeding
17 16 monomers exist at only insignificant concentrations. For injection of 320 nm monomers (not
18 shown), significant concentrations are found only for monomers and fractals composed of 2 and
19 4 monomers even with 8 Tg yr⁻¹ of emission.

20 Figure 5(a) shows the stratospheric alumina burden as a function of injection rate for four
21 different radii of injected monomers. Alumina burden is seen to be approximately linear with
22 injection rate. This is in contrast to a more strongly decreasing rate of change with increasing
23 injection rate seen for geoengineering by injection of SO₂ or H₂SO₄, also shown in Fig. 5(a). In
24 the case of sulfur injection, particles that grow to larger spherical sizes have shorter atmospheric
25 residence times. For alumina particles with sparse fractal structure ($D_f=1.6$), the fractal particles
26 do not increase their sedimentation velocities in the middle and upper stratosphere as they grow
27 by coagulation, resulting in residence times remaining almost constant over the alumina size
28 distribution. The cases that produce the fewest fractals ($R_0=240$ and 320 nm) have the most
29 linear response. The calculated atmospheric burden for diamond (not shown) is almost identical
30 to that for alumina of the same size injected monomer. Also shown in Fig. 5 as dashed lines are
31 simulations with coated alumina particles assumed to adopt a more compact fractal shape

1 ($D_f=2.8$). For these scenarios, total stratospheric burden is reduced due to the faster
2 sedimentation of the coated fractal particles, while the mass fraction in monomers is increased
3 due to fewer high-order fractals to scavenge the monomers. Only the 80 nm and 160 nm cases
4 show significant differences under the assumption that coated particles become more compact.

5 The stratospheric burden of sulfate is shown in Fig. 5(b) under various geoengineering scenarios
6 with alumina injection. Thick lines (left hand axis labels) represent the total stratospheric burden
7 of condensed sulfate as a function of geoengineering injection rate of alumina while thin lines
8 (right hand axis labels) represent the fraction of stratospheric liquid sulfate on the surface of
9 alumina particles. With injection of 80 nm alumina monomers, total stratospheric sulfate
10 increases above background for injections less than 1.5 Tg yr^{-1} , but then decreases with higher
11 injection rates. Up to 86% of the total stratospheric sulfate is found on alumina particles in these
12 cases, a result of the large alumina surface area available and high coagulation rates with large
13 fractals. Alumina injection cases with larger monomer diameters lead to decreases in the total
14 stratospheric burden of liquid sulfate because of the faster sedimentation of the larger alumina
15 particles along with their sulfate coatings. The maximum decrease in total stratospheric sulfate
16 is about 30%. The fraction of total stratospheric sulfate found on alumina particles is as much as
17 82% with 160 nm monomers, 61% with 240 nm monomers, and 32% with 320 nm monomers.
18 The calculated thickness of the sulfate coating on alumina particles in the stratosphere varies
19 from 5 to 15 nm with 80 nm monomers and from 10 to 40 nm with 240 nm monomers with 1 Tg
20 yr^{-1} of injection. However, as the geoengineering injection rate increases, the sulfate layer on
21 alumina particles becomes thinner since the stratospheric sulfate burden will then be distributed
22 over a larger alumina surface area.

23

24 **3.4 Radiative forcing assessment**

25 In this work, we confine our assessment of radiative forcing to reflected solar radiation,
26 acknowledging that absorption of thermal infrared radiation generates a longwave radiative
27 forcing as well. For some materials, this longwave forcing can be significant, reaching
28 approximately 20% of the shortwave radiative forcing in the case of sulfate. Compared to the
29 shortwave forcing, however, the longwave forcing is a much more sensitive function of profiles
30 of atmospheric temperature, humidity, ozone, and well-mixed greenhouse gases. Furthermore,

1 [the instantaneous longwave forcing adjusts in response to changes in the temperature profile](#)
2 [caused by radiative heating rate perturbations due to the aerosol particles. We therefore restrict](#)
3 [our analysis to shortwave radiative forcing to narrow the scope of radiative transfer calculations](#)
4 [and to reduce the number of radiative forcing values presented.](#)

5 Alumina particles are known to be more efficient scatterers than sulfate particles, and thus are
6 expected to be more efficient per unit mass for geoengineering applications. Figure 6 compares
7 the Mie scattering properties in the solar band of alumina and diamond monomers and sulfate
8 particles as a function of particle radius. We calculated the solid particle monodisperse single
9 scatter albedo values from Mie Theory (Bohren and Huffman 2008) using tabulated complex
10 refractive index data for diamond (Edwards and Philipp 1985) and alumina (Thomas and Tropic
11 1997). The upscatter and downscatter cross sections are calculated from Wiscombe and Grams
12 (1976), utilizing the scattering phase function from Mie Theory and the same complex refractive
13 index data. Figure 6(a) shows the ratio of upscatter cross section to geometric cross-section for
14 the three particle types. The sulfate profile is fairly flat, with a cross section of about 0.3 for
15 particles greater than 0.5 μm , whereas the alumina profile shows a 30% drop from its peak of 0.6
16 between 0.2 μm and 0.6 μm to 0.4 at 2 μm . The diamond profile shows a peak of 0.9 between
17 0.15 and 0.5 μm , dropping to about 0.55 at radii greater than 1.2 μm . Figure 6(b) shows strong
18 peaks in upscatter per unit volume for alumina and diamond monomers as a function of radius.
19 In contrast, sulfate particles exhibit a much flatter function of upscatter per unit volume as a
20 function of radius. Alumina monomers scatter most efficiently per unit particle volume at about
21 200-250 nm. At this radius, they have three times the upscatter per unit volume as sulfate
22 particles. Upscatter per unit mass however, due to the difference in density of alumina relative to
23 sulfate, shows less contrast. For diamond monomers, the upscatter per unit volume peaks at
24 around 150 nm radius, with over twice the peak upscatter of alumina monomers. Figure 6(c)
25 shows the ratio of downscatter cross section to upscatter cross section for alumina, diamond, and
26 sulfate as a function of radius. Alumina monomers have about half the downscatter per unit of
27 upscatter as sulfate particles, while diamond monomers have half the downscatter of alumina.
28 Thus in geoengineering applications, alumina and diamond would scatter radiation back to space
29 and produce substantially smaller increases in diffuse radiation at the surface than would sulfate
30 particles producing the same change in RF.

1 We use a scattering code which integrates the Mie scattering function over shortwave spectral
2 bands and scattering angles as a function of particle size (monomers and fractals) using an
3 efficient impulse-function method. Multiple scattering is ignored as solid particle optical depths
4 are small. For purposes of radiative forcing, we assume that solid particles with thin sulfate
5 coatings behave the same as bare particles. We follow Rannou et al. (1999) for scattering by
6 fractals, and follow the approximation in Charlson et al. (1991) by scaling our calculated RF
7 values by $(1-\alpha)^2$, where α represents surface albedo, here taken to be 0.2. Figure 7(a) shows the
8 shortwave globally-averaged clear-sky radiative forcing functions, in W m^{-2} per Tg of aerosol
9 burden, obtained by our scattering code as a function of monomer radius and fractal size (number
10 of cores per particle). Scattering by 80 nm alumina monomers is much less efficient (factor of 4)
11 than scattering by 160 nm alumina monomers. There is little difference in scattering between
12 240 nm and 320 nm alumina monomers, both with about 50% greater RF per terragram burden
13 than 160 nm monomers. Fractals scatter much less efficiently than monomers. A fractal
14 aggregate of two cores scatters only 50% as much radiation per unit mass as a corresponding
15 monomer, and higher order fractals scatter even less efficiently. An aggregate of 16 alumina
16 monomers has negligible scattering per unit mass. The functions with fractal dimension of both
17 $D_f=1.6$ (solid lines) and $D_f=2.8$ (dashed lines, labeled “compact coated”) are shown for
18 alumina, however, this produces only a minor difference in radiative forcing per unit mass. The
19 radiative forcing function for diamond with 160 nm monomers (the radius of peak backscatter
20 efficiency) shows significantly greater forcing than alumina, 2.7 times greater than 160 nm
21 alumina monomers and 1.8 times greater than 240 nm alumina monomers.

22 We obtain averages of solid aerosol mass in each bin size (monomers and fractals) integrated
23 vertically and averaged over latitude and season. The integrated and averaged aerosol mass per
24 bin is multiplied by the spectrally-integrated radiative forcing per terragram burden for each bin
25 to obtain the total radiative forcing for each geoengineering scenario. The global annual average
26 top-of-atmosphere shortwave radiative forcing due to alumina and diamond is shown in Fig. 7(b)
27 as a function of injection rate for specified sizes of injected monomers under clear sky
28 conditions. [As stated at the beginning of this section, our analysis focuses on shortwave
29 radiative forcing, as it provides for a less ambiguous comparison, relative to combined shortwave
30 plus longwave forcing, of the efficacy of different particles in offsetting surface warming.](#)

Debra Weisenstein 10/6/2015 11:55 AM

Deleted: We choose to present only shortwave radiative forcing as more relevant to geoengineering intended to offset surface warming after atmospheric adjustments.

1 | However, [we note that](#) the longwave radiative forcing is about 10% of the shortwave RF for
2 alumina, though of opposite sign, and is negligible for diamond.

3 For alumina, shortwave RF for cases with 80, 160, 240, and 320 nm injected monomer size are
4 shown in Fig. 7(b) as a function of injection rate. Cases with 80 nm monomer injections have
5 very low RF, due both to inefficient scattering for monomers of that size, and the large
6 proportion of fractals to monomers. The RF for the 80 nm injection case increases very little
7 with increasing injection rate, as increasing injections produce fractals composed of more than
8 64 monomers which produce almost no scattering per unit mass. The case with injection of 320
9 nm alumina monomers produces less RF than the case with injection of 240 nm alumina
10 monomers. Though monomers of 320 nm produce slightly more RF per terragram than
11 monomers of 240 nm, the 320 nm injection cases yield a smaller burden due to their faster
12 sedimentation rates. Injection of 240 nm monomers is found to produce the most radiative
13 forcing per terragram of alumina injected annually, consistent with the peak of the upscatter per
14 unit volume curve shown in Fig. 6(b). We calculate radiative forcing for diamond injections of
15 160 nm monomers only, also shown in Fig. 7(b). Atmospheric burden of diamond is very
16 similar to that for alumina of the same radius, but RF is much larger owing to more efficient
17 scattering. Diamond injection at a rate of 4 Tg yr^{-1} results in -1.8 W m^{-2} of shortwave forcing,
18 while the same alumina injection results in only -1.2 W m^{-2} of shortwave forcing. The increase
19 in downward diffusive flux is also calculated by our radiative forcing code and is shown in Table
20 2 for selected cases which each produce -2 W m^{-2} of shortwave radiative forcing. For equivalent
21 radii and injection rate, diamond produces up to twice the diffuse downward radiation as
22 alumina, however, per unit change in top-of-atmosphere shortwave radiative forcing, diamond
23 produces less diffuse downward radiation.

24 | Our method produces only a globally-averaged value of [shortwave](#) radiative forcing by solid
25 particles. Our results are not meant to be of high accuracy, as they do not account for clouds or
26 molecular scattering and are limited by the index of refraction data, uncertainties in fractal
27 scattering, and our averaging method. Nevertheless, it is useful to obtain well-founded estimates
28 of radiative forcing for comparison with sulfate geoengineering, and relative efficiencies among
29 solid particle scenarios as a function of injected monomer diameter. The RF plot in Fig. 7(b)
30 shows clear sky shortwave radiative forcing from two sulfur geoengineering scenarios as well.
31 The scenario results were calculated with the AER 2-D model, as applied in Pierce et al. (2010)

1 but using the radiative scattering code applied to alumina and diamond. Note that we plot them
2 here relative to the total SO₂ or H₂SO₄ injection mass per year, not the sulfur mass injected per
3 year. The most efficient alumina geoengineering scenario, with 240 nm monomers injected, has
4 roughly the same RF efficiency per terragram of injection as geoengineering by injection of
5 H₂SO₄. However, if a geoengineering methodology were to transport only sulfur to the
6 stratosphere and create H₂SO₄ in situ, then sulfur geoengineering would be more efficient than
7 alumina per terragram per year transported.

8 Aerosol heating of the tropical lower stratosphere is another potential risk of geoengineering.
9 Heckendorn et al. (2009) investigated this effect and the resulting increase in stratospheric water
10 vapor, primarily caused by longwave heating, for sulfate aerosol. To estimate lower
11 stratospheric heating by solid particles, we use the Rapid Radiative Transfer Model (RRTM)
12 developed by Atmospheric and Environmental Research (Mlawer et al. 1997; Clough et al. 2005)
13 to calculate radiative heating rates for mean cloud-free tropical atmospheric profiles with and
14 without a uniform aerosol density of 1 cm⁻³ between 18 and 23 km. The combined longwave and
15 shortwave heating rates shown in Table 2 for alumina, diamond, and sulfate are generated by
16 scaling the RRTM results for number densities of 1 cm⁻³ to the average number density in the 18-
17 23 km region between 30°S and 30°N for scenarios predicted to produce -2 W m⁻² of shortwave
18 radiative forcing. For alumina and diamond, the RRTM calculation uses only the monomer size
19 of 240 nm or 160 nm, respectively, ignoring fractal particles and treating coated monomers the
20 same as uncoated monomers. The range provided for heating by alumina and diamond in Table
21 2 uses monomer number densities for the low estimate and total particle number densities for the
22 high estimate as multipliers for the heating rate determined from RRTM with average number
23 density of 1 cm⁻³. For sulfate particles, we employ a size distribution due to the sensitivity of
24 heating rates to particle diameter and the range of diameters generated in geoengineering
25 scenarios. We find that the lower stratospheric heating rate from alumina is approximately 4-5
26 times less than the heating rate from sulfate, comparing scenarios which each generate -2 W m⁻²
27 of shortwave RF. Shortwave heating from alumina is about 15% of the total heating [by gases](#)
28 [and aerosols](#), and from sulfate about 20%. The total heating rate from diamond is almost entirely
29 due to shortwave effects, but is still much less than that for alumina with the same top-of-
30 atmosphere shortwave radiative forcing.

31

1 **3.5 Ozone impacts**

2 Heterogeneous reactions on stratospheric particles play an important role in ozone chemistry by
3 converting inactive forms of chlorine and bromine to forms that contribute directly to catalytic
4 destruction of ozone. In addition, the heterogeneous conversion of N_2O_5 to HNO_3 reduces NO_x
5 concentrations. This increases ozone concentrations in the middle stratosphere where NO_x
6 reactions dominate the ozone loss cycles, but it decreases ozone concentrations in the lower
7 stratosphere where HO_x , ClO_x , and BrO_x loss cycles dominate. Transient increases in sulfate
8 aerosols following volcanic eruptions have caused temporary depletions in ozone (Solomon,
9 1999). Geoengineering by stratospheric aerosol injection would be expected to lead to analogous
10 ozone depletion, depending on the heterogeneous reactions that occur on the particle surfaces
11 and their rates.

12 Ozone loss due to geoengineering injections of sulfate precursors has been explored by several
13 authors (Heckendorn et al., 2009; Tilmes et al., 2008; 2009; 2012; Pitari et al., 2014). Here we
14 provide a preliminary assessment of ozone loss from geoengineering injection of alumina and
15 diamond solid particles. To enable a relative comparison of the ozone impact of sulfate
16 geoengineering, we use the same model to compute changes in ozone abundance arising from
17 injections of both solid particles and of sulfate aerosols. We use the AER 2-D chemistry-
18 transport-aerosol model which includes full ozone chemistry, with 50 transported species, an
19 additional 51 radical species, 286 two- and three-body chemical reaction, 89 photolysis
20 reactions, and 16 rainout/washout removal processes coupled to our aerosol model (Weisenstein
21 et al., 1998; 2004; Rinsland et al., 2003). Reaction rates are from the JPL compendium (Sander
22 et al., 2011). The model parameterizes polar stratospheric clouds (PSCs) using thermodynamic
23 equilibrium, employing the formulas of Hanson and Mauersberger (1998) and Marti and
24 Mauersberger (1993) for equilibrium vapor pressures over solid HNO_3 and ice, respectively,
25 assuming no supersaturation and prescribing the particle radii. A comparison with observed
26 ozone trends between 1979 and 2000 is presented in Anderson et al. (2006) for the AER model
27 and several other models. Our simulations of ozone change due to SO_2 injections are similar to
28 those of Heckendorn et al. (2009) if we compare equivalent scenarios, but larger than those of
29 Tilmes et al. (2012). This model does not include radiative or dynamical feedbacks;
30 temperature and circulation are fixed with a climatology averaged over the years 1978 through
31 2004. Thus our evaluation of ozone changes due to geoengineering by injection of solid

1 particles includes only chemical perturbations due to heterogeneous reactions on particle
2 surfaces and not due to changes in temperature or dynamics induced by the geoengineering.
3 The amount of ozone loss induced by stratospheric aerosols is strongly dependent on the
4 concentrations of halogen species, primarily Cl and Br. Future concentrations of halogens are
5 expected to decline as a result of emissions controls, so the impact of stratospheric aerosols on
6 ozone will—all else being equal—decline over time. To err on the side of caution by overstating
7 the ozone impacts, we use present-day trace gas concentration throughout this study with 3.4
8 ppbv of total chlorine and 23 pptv of total bromine, including 6 pptv of inorganic bromine
9 derived from the very short lived (VSL) organic compounds CH_2Br_2 and CHBr_3 . Tilmes et al.
10 (2012) showed that inclusion of VSL bromine increases ozone depletion in geoengineering
11 scenarios. A more detailed evaluation of ozone impacts of solid particle geoengineering will
12 await further studies with coupled aerosol-chemistry-climate models.

13 Aerosol surface area density (SAD in units of $\mu\text{m}^2/\text{cm}^3$) contributes to determining the rates of
14 heterogeneous chemical reactions that occur on particle surfaces. Heterogeneous reactions may
15 occur on bare alumina surfaces in the stratosphere, as well as on sulfate surfaces. The reaction
16 $\text{ClONO}_2 + \text{HCl} \rightarrow \text{Cl}_2 + \text{HNO}_3$ has been measured on alumina surfaces (Molina et al., 1997;
17 Sander et al., 2011) and would be expected to cause ozone depletion (Danilin et al., 2001;
18 Jackman et al., 1998), though uncertainties in this reaction and the surface properties of alumina
19 aerosol remain unexplored. Figure 8 shows bare alumina surface area density for the cases with
20 1 Tg yr^{-1} injection of 80 nm monomers (left panel) and 240 nm monomers (right panel).
21 Alumina SAD is largest in the tropics where particles are injected, and is lower at higher
22 latitudes where a larger fraction of the surfaces are coated with sulfate. Alumina SAD extends to
23 higher altitudes, up to 40 km and above, with injection of 80 nm monomers, whereas the 240 nm
24 monomers and their fractal derivatives sediment fast enough to preclude significant SAD above
25 30 km. Note that sulfate aerosols generally evaporate above about 35 km altitude, and thus
26 geoengineering with solid particles may introduce significant surface area density in regions that
27 currently are not greatly impacted by heterogeneous chemistry.

28 Figure 9 shows sulfate SAD from the calculated background atmosphere without geoengineering
29 (panel a) and the increase in sulfate SAD in an atmosphere with 1 Tg yr^{-1} of geoengineering
30 injection of 80 nm alumina monomers (panel b) or 240 nm alumina monomers (panel c). While

1 the addition of alumina particles has produced only a small change in the total stratospheric
2 condensed sulfate (see Fig. 5b), it has produced significant increases in sulfate surface area
3 density. This is a result of sulfate being distributed in thin layers on the surfaces of alumina
4 particles. With injection of 80 nm monomers, the sulfate SAD has increased by factors of 2-4 in
5 the lower stratosphere, with maximum SAD at high latitudes where significant concentration of
6 complex alumina fractals exist to scavenge the smaller sulfate particles. With injection of 240
7 nm monomers, the maximum sulfate SAD occurs in the tropics as the faster sedimentation of
8 alumina in this case results in a smaller concentration of mostly monomers at high latitudes.
9 Figure 9 (panel d) shows the increase in sulfate SAD for a geoengineering scenario with 1 Tg
10 yr^{-1} of SO_2 injection as calculated by the AER 2-D model. The SAD increase is similar in
11 magnitude to the case with injection of 80 nm alumina monomers, but has a distribution similar
12 to the 240 nm alumina injection case. A similar injection of sulfur as H_2SO_4 , as in Pierce et al.
13 (2010), produces more than double this SAD increase. For reactions whose rate is dominated by
14 liquid sulfate surface area density, we would expect similar chemical ozone loss from similar
15 changes in sulfate SAD whether due to geoengineering by SO_2 , H_2SO_4 , or alumina injection.

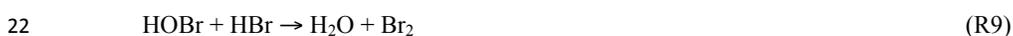
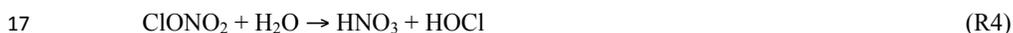
16 The SAD generated by alumina geoengineering is reduced when the monomer size of the
17 injected particles increases. Optimizing the injected monomer size would be an important
18 strategy to minimize stratospheric ozone depletion. Figure 10 illustrates this, showing annual
19 averaged SAD between 15 and 25 km for bare alumina (panel a) and total sulfate (pure sulfate
20 plus sulfate-coated alumina, panel b) as a function of injection rate with injections of 80, 160,
21 240, and 320 nm monomers. The SAD for bare alumina drops by factors of 1.8 to 3.1,
22 depending on injection rate, when the monomer size is increased from 80 nm to 160 nm. The
23 alumina SAD is roughly linear with injection rate, since the alumina surface area density doesn't
24 decrease as particles coagulate when $D_h=2.0$. The total sulfate SAD (Fig. 10b) is even more
25 dependent on monomer diameter than is the uncoated alumina SAD. Even though the burden of
26 stratospheric sulfate on alumina varies slowly with injection rate, the sulfate SAD varies more
27 rapidly with injection rate as the sulfate becomes spread more thinly over a greater numbers of
28 alumina particles. The dashed lines in Fig. 10 represent cases where the coated particles take on
29 a more compact fractal shape, and thus sediment faster, greatly decreasing sulfate SAD for the
30 80 and 160 nm injection case. Our diamond simulation is similar to the alumina simulation with
31 injection of 160 nm monomers.

1 Our model evaluation of ozone impacts from alumina geoengineering considers the following
2 reaction on bare alumina surfaces:



4 This reaction, with reaction probability γ of 0.02, has been studied in relation to ozone depletion
5 resulting from space shuttle launches (Danilin et al., 2001; Jackman et al., 1998). We assume
6 that this reaction occurs catalytically with no surface poisoning. Other potential reactions on
7 alumina surfaces have been investigated in the laboratory (see Sander et al., 2011), and further
8 investigation is needed to determine how additional heterogeneous and photocatalytic reactions
9 could modify stratospheric chemistry or change the surface properties of alumina in the
10 stratosphere. Measurements of potential heterogeneous reactions on diamond surfaces are not
11 available. Thus we perform diamond injection calculations assuming (R1) on diamond at the
12 same rate as for alumina, and assuming no reactions on bare diamond surfaces.

13 We also consider heterogeneous reactions on the sulfate-coated surfaces of solid particles. These
14 reactions include:



23 Due to the solubility of HCl, ClONO₂, HOCl, and HOBr in sulfuric acid solutions, reactions
24 (R3), (R4), (R5), and (R8) can be considered bulk processes or hybrid bulk-surface processes
25 governed by a reaction-diffusion process. For liquid spherical particles we use standard methods
26 to calculating the reaction probability as a function of radius (Shi et al., 2001). For reactions on

1 coated solid particles, we use the same functions, substituting the thickness of the sulfate layer in
2 place of the radius of a spherical particle.

3 Calculated changes in ozone due to heterogeneous reactions on alumina surfaces (bare alumina
4 and sulfate-coated alumina) are shown in Fig. 11 for cases with injection of 80 nm and 240 nm
5 monomers at an injection rate of 1 Tg yr^{-1} . Column ozone depletion ranges from 2% in the
6 tropics to 6-10% at midlatitudes and up to 14% at the poles in springtime with injection of 80 nm
7 monomers. With injection of 240 nm monomers, ozone depletion is much smaller, ranging from
8 0.3% to 2.5%. The annual average ozone change as a function of latitude and altitude (Fig. 11,
9 right hand panels) shows features linked to local balances in ozone's formation rate, chemical
10 destruction rate, and local transport rates. In the tropics, ozone concentrations are determined by
11 the chemical production via UV radiation that is balanced by transport out of the tropics to
12 higher latitudes. Thus ozone changes due to increased loss mechanisms are minimal in the
13 tropics in the stratosphere, though increased penetration of UV to lower altitudes in the tropics
14 can produce ozone increases. In the middle stratosphere at 25-35 km altitude, the NO_x cycle
15 dominates ozone loss. Increases in aerosol surface area density in this region reduce NO and
16 NO_2 while increasing HNO_3 via the $\text{N}_2\text{O}_5 + \text{H}_2\text{O}$ reaction. Thus the NO_x loss cycle is diminished
17 and ozone increases. The sedimentation rate of alumina particles is significant, as the scenario
18 with injection of 80 nm monomers yields significant increases in aerosol surface area density and
19 ozone changes above 25 km, whereas the scenario with 240 nm monomers injected does not. In
20 the lower stratosphere at mid and high latitudes, the heterogeneous reactions on particle surfaces
21 increase the ratio of chlorine and bromine in their radical forms that destroy ozone (Cl, ClO, Br,
22 and BrO). In addition, the $\text{N}_2\text{O}_5 + \text{H}_2\text{O}$ reaction leads to more HNO_3 and less ClONO₂ and
23 BrONO₂, thus indirectly increasing halogen radicals as well. Local ozone depletions in the lower
24 stratosphere are as large as 24% with 80 nm monomer injections and 5% with 240 nm monomer
25 injections, on an annual average basis.

26 Figure 12(a) shows annual average changes in ozone column as a function of latitude with 1 Tg
27 yr^{-1} of geoengineering injections. Results with injection of 80 nm, 160 nm, and 240 nm alumina
28 monomers are shown, along with injection of 160 nm diamond monomers. We don't calculate
29 ozone changes due to injection of 320 nm alumina monomers because these scenarios produce
30 less radiative forcing than injection of 240 nm monomers for similar injection rates. Ozone
31 changes, similar to SAD increases, are found to be very sensitive to injected monomer size.

1 However, assuming that coated alumina particles assume a more compact shape (shown by
2 dashed lines in the figure) significantly reduces calculated ozone depletion for the $R_0=80$ nm
3 case, and modestly reduces ozone depletion for the $R_0=160$ nm case. Also shown in Fig. 12(a)
4 are changes in ozone column due to injection of 1 Tg yr^{-1} of SO_2 and H_2SO_4 , which are roughly
5 similar to those calculated with 1 Tg yr^{-1} of 80 nm alumina monomers.

6 Figure 12(b) shows calculated ozone changes for injection rates of 1, 2, 4, and 8 Tg yr^{-1} with
7 injection of 240 nm alumina monomers. Note that ozone changes increase at a less than linear
8 rate with increasing injection rates, and that the effect of compaction of coated alumina particles
9 becomes more significant at higher injection rates due to formation of higher-order fractals.

10 Figure 12(c) shows calculated ozone changes due to geoengineering injection of diamond
11 monomers of 160 nm radius. Solid lines are for results including the (R1) reaction on uncoated
12 particles, and dotted lines omit this reaction. Reaction (R1) has a greater effect in the tropics
13 than at mid latitudes due to higher concentrations of uncoated particles there. The northern high
14 latitudes show greater sensitivity to geoengineering injections than the southern high latitudes at
15 the higher emission levels, likely due to the dominant role of PSCs over the Antarctic.

16 Global average column ozone changes are shown in Fig. 13(a) as functions of injected monomer
17 size and injection rate. Figure 13(b) shows changes in global average ozone as a function of the
18 associated radiative forcing for each scenario. This makes it clear that geoengineering injection
19 of 80 nm alumina monomers is completely unworkable, producing excessive ozone depletion
20 (5% with 1 Tg yr^{-1} injection and 14% with 8 Tg yr^{-1} injection) and minimal radiative forcing.
21 Geoengineering by injection of 240 nm alumina monomers, however, could potentially be an
22 effective climate control strategy, similar to geoengineering by injection of sulfur in its radiative
23 forcing effectiveness but with less ozone depletion potential. Note that radiative forcing and
24 associated ozone depletion with 16 Tg yr^{-1} injection of 240 nm alumina monomers is included in
25 the Fig. 13(b), yielding 3.6 W m^{-2} of RF with about 8% ozone depletion. Injection of 160 nm
26 diamond monomers produces ozone loss per terragram of injection similar to 160 nm alumina
27 monomers, but with radiative forcing per terragram of injection greater than for similar injections
28 of 240 nm alumina monomers. We show diamond results both including and excluding reaction
29 (R1) on bare diamond surfaces. This reaction makes about a 10-15% difference in ozone
30 depletion due to diamond injection. Note that the SO_2 results plotted in Fig.13 show more ozone
31 depletion than in Heckendorn et al. (2009) because that study did not include the short-lived

1 bromine source gases and used a narrower injection region. More studies will be needed to
2 evaluate potential impacts on stratospheric chemistry such as tropopause heating and changes in
3 the Brewer-Dobson circulation that are not evaluated here.

4

5 **4 Discussion**

6 We have developed a new aerosol model and used it to quantitatively explore the interactions of
7 solid particles with sulfate aerosol in the stratosphere. This analysis allows a preliminary
8 assessment of some of the trade-offs that might arise in using solid aerosols such as alumina or
9 diamond rather than sulfates for solar geoengineering. We first discuss salient limitations of our
10 model before turning to analysis of trade-offs.

11 **4.1 Limitations**

12 *Injection mechanism.* We do not model the mechanism for injection and dispersion of aerosols. If
13 aerosols were injected from aircraft there would be small-scale dynamical effects in the injection
14 nozzle and in the aircraft plume in which particle concentrations would be much larger than
15 found after dilution to the scale of a model grid box, possibly leading to rapid coagulation.
16 Effects during particle generation or injection from a nozzle would occur on very short time and
17 space scales and cannot be estimated here. We can, however, estimate the impact of coagulation
18 in an expanding plume using the method of Pierce et al. (2010). We allow the plume cross-
19 section to dilute from 6 m^2 to $17 \times 10^6 \text{ m}^2$ over a 48 hours period, assuming that alumina particles
20 are release at a rate of 30 kg km^{-1} of flight path. We find the fraction of alumina mass remaining
21 as monomers after 48 hours of plume dilution to be 37%, 86%, 96%, and 98% for injected
22 monomers of 80, 160, 240, and 320 nm, respectively. For monomer injections of 240 and 320
23 nm, only 2-monomer fractals are created within 48 hours. We conclude that plume dynamics
24 and processing are unlikely to have a substantial effect on alumina geoengineering if injected
25 monomer size is greater than 160 nm. For 240 nm monomer, the most relevant case, our 2-D
26 model calculation would be expected to have 4% less mass in monomers if plume dynamics
27 were considered.

28 *Two-dimensional model.* A second limitation is the use of a 2-D model. Since the
29 geoengineering scenarios discussed in this work deal with particle injection in the 20-25 km

1 altitude region and spread between 30°S and 30°N, assuming zonal symmetry, as a 2-D model
2 implicitly does, does not detract greatly from the validity of our results. In particular, if the
3 method of injecting alumina particles attempts to distribute them uniformly in space and time
4 and avoid overlapping emissions as much as possible, then a zonally-symmetric spread may be a
5 fairly good approximation. However, details of transport near and below the tropopause are not
6 well-represented in 2-D models. Thus a 3-D model would be needed to accurately represent this
7 region. And if a specific geoengineering injection methodology were to be investigated, a 3-D
8 model with fine resolution would be needed to examine heterogeneities in the resulting aerosol
9 distribution.

10 *Geometry of aggregates, effects of size binning.* The fractal geometry of aggregates likely
11 depends on the formation mechanism, and it is plausible that the actual fractal dimensions might
12 differ significantly from the D_f value of 1.6 we use here for alumina and diamond. While the
13 fractal dimension of alumina has been measured for monomer cores much smaller than
14 considered here, that of diamond has not. It is also plausible that variables k_f and k_h should have
15 values other than 1.0, at least for cases that calculate many high-order fractals. The behavior of
16 aggregates under stratospheric conditions has not been studied extensively. The formulations we
17 have adopted for coagulation, condensation, and sedimentation are based on theoretical studies
18 or on tropospheric or liquid-medium experiments, and thus should be considered uncertain. Our
19 formulation also assumes that all injected monomers are of a uniform radius. While it appears
20 reasonable to assume that industrial production of alumina or diamond nanoparticles could
21 produce particles within a narrow size range, our assumption is a simplification. Likewise, our
22 assumption of maximal compaction instantaneously on wetting is likely not realistic but meant to
23 show the greatest possible affect of potential particle compaction on aging. Observational
24 studies in the laboratory and in the stratosphere would be needed to determine whether
25 compaction of alumina particles occurs and to what extent. However, compaction has a minor
26 effect on the radiative properties and ozone depletion potential of particles with monomer sizes
27 of ~200 nm or greater.

28 Numerical errors result from the discrete aerosol size binning we employ. The discretization
29 leads to a broadening of the size distribution during the coagulation process. Appendix A details
30 the coagulation methodology, in which coagulation of two solid particles often leads to a new
31 particle with size intermediate between two bins. In this case, particle mass is apportioned

1 between two bins, leading to mass transfer into a bin larger than that of the combined new
2 particle. This broadening of the distribution will lead to somewhat excessive sedimentation,
3 whose error depends on the coarseness of the bin spacing. Coagulation between liquid and solid
4 particles does not produce numerical broadening of the size distribution as the binning for mixed
5 solid-liquid particles depends only on the size of the solid particle.

6 *Ozone chemistry.* The surface chemistry of alumina and other solid particles potentially useful
7 for geoengineering has not been studied as extensively as that of sulfate particles. We include
8 only one reaction, $\text{ClONO}_2 + \text{HCl}$, on alumina and diamond particles in this modeling study.
9 Laboratory studies have investigated some additional reactions on alumina surfaces, and there
10 may be others not yet explored. Reported reactions on Al_2O_3 surfaces include the uptake of NO_2
11 and HNO_3 and reactions of several volatile organic compounds, including formaldehyde,
12 methanol, and acetic acid (Sander et al., 2011). In addition, photocatalysis reactions of several
13 species on Al_2O_3 surfaces have been reported (de Richter and Caillol, 2011), and may depend on
14 the exact composition or impurities of the particle surface. Photocatalysis of CFC compounds
15 has been considered as a method to mitigate the atmospheric burden of greenhouse gases if
16 augmented by artificial UV radiation in the troposphere. However, if these reactions were
17 effective in the stratosphere, they would contribute to the formation of free radical chlorine and
18 bromine, possibly increasing ozone depletion while reducing the lifetime the CFCs. Studies of
19 these and other reactions under stratospheric photochemical conditions would need to be
20 performed on any solid particle under consideration for geoengineering application.

21 *Missing feedbacks.* The modeling we present utilizes temperature and transport fields uncoupled
22 from the model's chemistry and aerosols and is therefore missing a number of feedback
23 processes that may be important in the atmosphere and may significantly change the radiative
24 forcing or ozone depletion estimates given here. These include changes in stratospheric
25 temperature due to aerosol heating, which would modify rates of reactions important to ozone
26 formation and loss. Aerosol heating and enhanced equator-pole temperature gradients would
27 also modify the strength of the Brewer-Dobson circulation and the polar vortex, with impacts on
28 aerosol concentration, PSC formation, and ozone concentration. Increases in the temperature of
29 the tropical tropopause layer would increase the transport of water vapor across the tropopause,
30 increasing stratospheric H_2O and OH concentrations, and reducing ozone (Heckendorn et al.,
31 2009). These additional ozone changes would further modify stratospheric temperature and

1 circulation. However, Heckendorn et al. (2009) found that ozone loss due to heterogeneous
2 chemistry, without the dynamical effects of changes in temperature, water vapor, and the
3 Brewer-Dobson strength, accounted for 75% of the ozone change.

4 A more uncertain feedback process is the effect of enhanced aerosol concentrations on upper
5 tropospheric cloudiness and cloud radiative properties (Kuebbeler et al., 2012; Cirisan et al.,
6 2013). A general circulation model with stratospheric chemistry and aerosol and cloud
7 microphysics would be needed to evaluate these feedback effects.

8

9 **4.2 Principal findings**

10 Use of alumina particles for SRM is potentially useful only if the size of the injected monomers
11 is larger than about 150 nm; the best results are only seen if the monomer radius exceeds about
12 200 nm. The strong dependence on monomer size can be understood if one assumes that the
13 injection rate will be adjusted so as to produce a given radiative forcing, for example 2 W m^{-2} .
14 For alumina, the peak mass-specific upscattering efficiency occurs at a radius of $\sim 200 \text{ nm}$ (see
15 Fig. 6b). As the monomer size gets smaller a higher monomer density and mass injection rate is
16 required to maintain the specified radiative forcing. The coagulation rate increases as the square
17 of monomer density, so the fraction of monomers in aggregates increase rapidly with monomer
18 density. Finally, the mass-specific radiative forcing for aggregates decreases quickly with the
19 number of monomers in an aggregate (see Fig 7a), so the injection rate must be increased further
20 to maintain a fixed radiative forcing. The net effect is that the radiative efficacy, the global
21 radiative forcing per unit mass injection rate, declines very rapidly for particle radii below 150
22 nm. We find that alumina monomers with radii of roughly 240 nm provide the most radiative
23 forcing for a given injection rate. For particle sizes beyond 240 nm, the scattering efficiency
24 remains roughly constant while the sedimentation rate increases, contributing to a decrease in
25 radiative forcing efficiency per unit injection rate.

26 As a specific example, consider the injection of 240 nm alumina monomers at a rate of 4 Tg yr^{-1}
27 evenly distributed between 30°S to 30°N and from 20-25 km in altitude. This produces a
28 stratospheric burden of 4.6 Tg (see Fig. 5a) and global radiative forcing of 1.2 W m^{-2} (see Fig.
29 7b). Under these conditions, coagulation of alumina particles is minimal: 81% of the alumina is
30 in monomers and only 4% is in aggregates of more than two monomers (see Fig. 4c). Particle

1 densities are a maximum in the lower tropical stratosphere with peak concentrations of about 4
2 cm^{-3} . The net effect of interaction with the background stratospheric sulfate is that about 50% of
3 the stratospheric sulfate is found as a coating (Fig. 5b, right axis) with a typical depth of order 10
4 nm on the alumina particles. The total sulfate burden is reduced from 0.11 to 0.08 Tg (Fig. 5b,
5 left axis) because the relatively fast fall speeds of the alumina aerosol provide a sedimentation
6 sink for sulfates, yet the sulfate surface area density is increased by an average of $2 \mu\text{m}^2 \text{cm}^{-3}$ in
7 the lower stratosphere (Fig. 10b). The annual global average ozone column is reduced by 3.7%
8 (Fig. 13a) with maximum ozone loss of 4 to 7% over polar regions for this scenario and the
9 given modeling assumptions.

10 As with sulfate aerosols, ozone concentrations increase at altitudes around 30 km in the mid
11 stratosphere where the NO_x cycle dominates but this is more than compensated by the halogen-
12 catalyzed ozone loss in the lower stratosphere. And with injection of 240 nm monomers,
13 sedimentation is rapid enough to preclude significant aerosol concentrations above 25-30 km.
14 Most of the ozone impact of alumina aerosols is found to be due the increase in sulfate surface
15 area and heterogeneous reactions on the liquid sulfuric acid. This is because most of the alumina
16 particles are coated with sulfate at mid and high latitudes where ozone loss reactions largely
17 determine the ozone concentration. If the rate of (R1) is set to zero in our simulations, the
18 column ozone depletion changes by less than 15% in the extra-tropics, but up to 35% in the
19 tropics with injection 4 Tg yr^{-1} of 240 nm alumina monomers. Thus uncertainty in the rate of
20 (R1) or the nature of the uncoated alumina surface does not have a strong influence on our
21 calculated ozone impacts. If we assume that the alumina particle surfaces remain uncoated and
22 that (R1) occurs on all alumina particles, we find that the ozone depletion is much less than that
23 obtained when the surfaces do become coated, implying that (R1) on alumina surfaces has less
24 effect on ozone than do sulfate heterogeneous reactions on the same surface area, mostly due to
25 the effectiveness of heterogeneous bromine reactions.

26 We can achieve a similar radiative forcing of 1.2 W m^{-2} with injection of 160 nm radius diamond
27 monomers at 2 Tg yr^{-1} . This injection rate produces a stratospheric burden of 3.3 Tg of diamond.
28 The corresponding ozone depletion due to diamond injection ranges between 3.8% globally due
29 to increased sulfate surface area alone to 4.3% when we assume that (R1) occurs on the bare
30 surface of diamond particles with the same reaction rate employed for alumina. However, levels

1 of ozone depletion are highly uncertain as this reaction, and other potential heterogeneous
2 reactions on diamond surfaces, have not been measured.

3

4 **4.3 Comparison with sulfate aerosols**

5 Whatever method is used to create an artificial radiative forcing, solar geoengineering is—at
6 best—an imperfect method for reducing climate impacts. Any technology for producing radiative
7 forcing will have a set of *technology-specific impacts*, such as ozone loss arising from the
8 introduction of aerosol particles into the stratosphere. However the radiative forcing is produced,
9 the *efficacy* of SRM is inherently limited by the fact that a change in solar radiative forcing
10 cannot perfectly compensate for the radiative forcing caused by increasing greenhouse gases
11 (Kravitz et al., 2014; Curry et al., 2014). A central motivation for considering solid aerosols
12 rather than sulfates is that they might have less severe technology-specific risks. As discussed in
13 the introduction, the principle technology-specific risks or side-effects of sulfate aerosols are
14 ozone loss, increased diffuse light, and stratospheric heating.

15 Loss of stratospheric ozone and an increase in diffuse light have direct impacts on ecosystems
16 and human health. The consequences of stratospheric heating are indirect and more speculative.
17 Heating of the tropical tropopause layer (TTL) might be expected to increase the amount of water
18 vapor entering the stratosphere. An increase in TTL temperature of 1 K increases the
19 concentration of water vapor entering the stratosphere by about 0.8 ppmv (Kirk-Davidoff, 1999).
20 Geoengineering with sulfate aerosols might heat the TTL region by several degrees, increasing
21 stratospheric water vapor concentrations by more than 2 ppmv (Heckendorn et al., 2009). This
22 would in turn exacerbate ozone loss and create a positive radiative forcing that would offset
23 some of the reduction in forcing from SRM. While there is uncertainty about the exact
24 consequences of heating the lower stratosphere, it's reasonably certain that all-else-equal, a
25 geoengineering method that does not heat the low stratosphere is preferable to one that does.

26 We estimate stratospheric heating for alumina, diamond, and sulfate geoengineering scenarios
27 with the RRTM model, as described in Section 3.4. Our results for alumina are broadly
28 consistent with the results of Ferraro et al. (2011) for titania. Note, however, that Ross and
29 Shaeffer (2014) conclude that the positive infrared radiative forcing from alumina can be larger
30 than the negative radiative forcing from solar scattering by the same particles. We suspect that

1 part of this discrepancy comes from the fact that Ferraro et al. (2011) and this paper assume a
2 narrow size distribution close to the optimal for solar scattering, whereas Ross and Shaeffer
3 (2004) use a broad alumina size distribution. However, we have not resolved this discrepancy, so
4 our estimate of heating for alumina should be taken as uncertain.

5 As shown in Table 2, our results suggest that alumina may have less severe technology-specific
6 risks than sulfates. While the injected mass necessary to achieve a -2 W m^{-2} radiative forcing is
7 roughly equivalent whether employing alumina or sulfate aerosol, the ozone depletion is more
8 severe with sulfate geoengineering. In addition, the increase in diffuse solar radiation would be
9 half as much with alumina as with sulfate, and the stratospheric heating is expected to be
10 considerably less, smaller by a factor of 4-5 in our estimation. Diamond appears to offer
11 excellent shortwave scattering with only a small increase in diffuse light. We estimate ozone
12 depletion due to diamond to be less than that due to sulfate, but uncertainty is large. Lower
13 stratospheric heating from diamond is quite small.

14 We conclude that SRM by injection of solid particles may have some advantages relative to
15 sulfates and merits further study to reduce the sizable uncertainties that currently exist. It is
16 important to note that the injection of substances like alumina or diamond nanoparticles have
17 much greater “unknown unknowns” than sulfates, as they would be novel substances in the
18 stratosphere. Laboratory studies of reaction kinetics for these particles under stratospheric
19 conditions, as well as studies of their microphysical and radiative properties, are required to
20 reduce uncertainties.

21

22 **Appendix A: Microphysical schemes for coagulation and condensation**

23 A.1 Coagulation

24 A.1.1 Coagulation between particles of similar composition

25 In our sectional aerosol representation, collisions between particles of similar composition in
26 bins i and j , where bin j is smaller than bin i , produce particles of size intermediate between bins
27 i and $i+1$. Since the volume ratio between size bins, V_{rat} , is 2.0 in our case, coagulation between
28 two particles in bin i yields one particle in bin $i+1$. The rate of coagulation between particles in

1 bin i and particles in bin j , β_{ij} , is a function of the coagulation kernel, K_{ij} , and the number
 2 densities of the particles in each bin, N_i and N_j .

$$\beta_{ij} = SN_jN_iK_{ij}$$

3 where $S=0.5$ if $i=j$, otherwise $S=1.0$. A coagulation event between a particle in bin i and a
 4 particle in bin j ($j \leq i$) changes particle number densities in three bins as follows:

$$\frac{dN_j}{dt} = -\beta_{ij}$$

$$\frac{dN_i}{dt} = -\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

$$\frac{dN_{i+1}}{dt} = +\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

5 where $F = \frac{V_j}{V_i}$.

6 Coagulation of sulfate particles with other sulfate particles and of alumina particles with other
 7 alumina particles is generated by a double sum of the above equations over all bins i and j , where
 8 $j \leq i$. Our explicit formulation generates $\frac{dN}{dt}$ for all size bins before updating any values of N .

9 A.1.2 Coagulation between sulfate particles and solid particles

10 Coagulation between pure sulfate particles and pure alumina particles transfers particles from the
 11 category of pure alumina to the category of coated alumina, and shifts sulfate mass from pure
 12 sulfate particles to coated alumina sulfate mass. Since coated alumina particles are binned by
 13 their alumina mass alone, the alumina size index is unchanged in this coagulation process. For
 14 coagulations between an alumina particle of bin i and a sulfate particle of bin j , the rate of
 15 change of concentrations is:

$$\frac{dN_j(S)}{dt} = -\beta_{ij}$$

$$\frac{dN_i(Al)}{dt} = -\beta_{ij}$$

$$\frac{dN_i(c-Al)}{dt} = +\beta_{ij}$$

$$\frac{dM_i(S-Al)}{dt} = +\beta_{ij} DS_j$$

- 1 where $N(S)$, $N(Al)$, $N(c-Al)$ represent number concentrations of pure sulfate, pure alumina, and
 2 coated alumina particles, respectively, $M(S-Al)$ represents the mass of sulfate on a coated
 3 alumina particle, and DS_j represents the mass of sulfate in a single pure sulfate particle of bin j .
 4 Coagulation between a pure sulfate particle of bin j and a coated alumina particle of bin i
 5 removes a sulfate particle and transfers its sulfate mass onto the coated alumina particle:

$$\frac{dN_j(S)}{dt} = -\beta_{ij}$$

$$\frac{dM_i(S-Al)}{dt} = +\beta_{ij} DS_j$$

- 6 A.1.3 Coagulation between different types of solid particles
 7 For coagulation between an uncoated and a coated alumina particle:

$$\frac{dN_j(Al)}{dt} = -\beta_{ij}$$

$$\frac{dN_i(c-Al)}{dt} = -\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

$$\frac{dN_{i+1}(c-Al)}{dt} = +\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

$$\frac{dM_i(S-Al)}{dt} = -\beta_{ij} \frac{F}{(V_{rat} - 1)} DSAl_i$$

$$\frac{dM_{i+1}(S-Al)}{dt} = +\beta_{ij} \frac{F}{(V_{rat} - 1)} DSAl_i$$

- 8 where $DSAl_i$ represents the mass of sulfate on a single coated alumina particle of bin i , obtained
 9 as the ratio of total sulfate mass on solid particles of bin i for a given grid box to the number of
 10 coated alumina particles in that grid box.
 11 For coagulation between two coated alumina particles:

$$\frac{dN_j(c-Al)}{dt} = -\beta_{ij}$$

$$\frac{dN_i(c-Al)}{dt} = -\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

$$\frac{dN_{i+1}(c-Al)}{dt} = +\beta_{ij} \frac{F}{(V_{rat} - 1)}$$

$$\frac{dM_j(S-Al)}{dt} = -\beta_{ij} DSAL_j$$

$$\frac{dM_i(S-Al)}{dt} = -\beta_{ij} DSAL_i + \beta_{ij} \left(1 - \frac{F}{(V_{rat} - 1)}\right) (DSAL_j + DSAL_i)$$

$$\frac{dM_{i+1}(S-Al)}{dt} = +\beta_{ij} \frac{F}{(V_{rat} - 1)} (DSAL_j + DSAL_i)$$

1 A.1.4 Coagulation kernel

2 The coagulation kernels, K_{ij} , are calculated following classical theory as detailed in Jacobson
3 (1999). We apply only Brownian coagulation without convective, gravitational, or Van der
4 Waals corrections.

$$K_{ij} = \frac{4\pi(r_i + r_j)(D_i + D_j)}{\frac{r_i + r_j}{r_i + r_j + (\delta_i^2 + \delta_j^2)^{1/2}} + \frac{4(D_i + D_j)}{(r_i + r_j)(\bar{c}_i^2 + \bar{c}_j^2)^{1/2}}}$$

$$\delta_i = \frac{(2r_i + \lambda_i)^3 - (4r_i^2 + \lambda_i^2)^{3/2}}{6r_i\lambda_i} - 2r_i$$

5 where r_i represents particle radius, D_i represents the particle diffusion coefficient, \bar{c}_i represents
6 the thermal velocity of the particle, and λ_i represents the particle mean free path.

7 For liquid particles, we use the spherical radius in the calculation of coagulation kernels; for
8 solid particles we use the radius of gyration R_g in the calculation (Maricq, 2007), as this more
9 closely represents the distance over which fractal particles may interact with other particles.

10 A.2 Condensational Growth and Evaporation

11 The rate of H_2SO_4 vapor condensation onto sulfate particles or sulfate-coated alumina particles
12 of bin i follows Jacobson (1999) and is given by:

$$\frac{dN_{H_2SO_4}}{dt} = \frac{-S_i N_i D_v (N_{H_2SO_4} - \alpha N_{H_2SO_4}^{eq})}{r_i}$$

1 where S_i represents the surface area of a single particle in bin i , N_i the particle number density,
 2 D_v is the molecular diffusion coefficient of the H_2SO_4 vapor, $N_{H_2SO_4}$ represents the number
 3 density of ambient H_2SO_4 gas, and $N_{H_2SO_4}^{eq}$ the equilibrium H_2SO_4 number density above a liquid
 4 H_2SO_4 surface. α is the Kelvin effect correction term, representing the saturation vapor pressure
 5 over a curved surface relative to that over a flat surface, given by:

$$\alpha = \exp\left(\frac{2\sigma_p\mu_p}{r_i R^* T \rho_p}\right).$$

6 α is a function of the surface tension of the H_2SO_4 liquid, σ_p , the average molecular weight of
 7 the liquid, μ_p , the particle radius, r_i , gas constant R^* , temperature, T , and liquid solution density
 8 ρ_p . The diffusion coefficient is defined as:

$$D_v = \frac{3}{8d^2 N_{air}} \left(\frac{k_B T A_v}{2\pi\mu_{H_2SO_4}} \frac{\mu_{air} + \mu_{H_2SO_4}}{\mu_{air}} \right)^{0.5} \omega$$

$$\omega = \left\{ 1 + \left[\frac{1.3333 + 0.71K_n^{-1}}{1 + K_n^{-1}} \right] K_n \right\}^{-1}$$

9 where d is the diameter of a gas molecule, k_B is Boltzman's constant, and A_v is Avogadro's
 10 number, ω is a correction term for collision geometry, and K_n is the Knudsen number. We
 11 assume an accommodation coefficient of 1.0. We use the spherical radius in the calculations for
 12 pure sulfuric acid-water particles, and the radius of gyration, R_g , increased by the coating
 13 thickness, for the sulfate-coated solid particles. However, for the curvature term α , we use the
 14 radius of the monomers composing the fractal particles increased by the coating thickness. The
 15 particle surface area for fractal particles is as described in Section 2 of this paper.

16 In the case of pure liquid H_2SO_4 - H_2O particles, condensation increases the particle radius and
 17 shifts mass into larger size bins, much as the coagulation formulation described above distributes
 18 mass between two bins when a particle's radius would fall between two bins. For sulfate-coated
 19 alumina particles, the bin is defined by the radius of the alumina particle only, so the size bin
 20 remains constant while the mass of sulfate associated with that bin increases. Evaporation
 21 occurs via the identical equations when $N_{H_2SO_4}$ is less than $\alpha N_{H_2SO_4}^{eq}$ and $\frac{dN_{H_2SO_4}}{dt}$ becomes
 22 positive. This occurs above about 35 km for sulfuric acid particles.

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Author contributions

D. Keith conceived this study and provided overall scientific direction. D. Weisenstein developed the solid-liquid aerosol model and performed the 2-D calculations. D. Keith performed the SW radiative forcing calculations, while J. Dykema calculated the heating rates with LW and SW contributions.

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28

1 Table 1. Overview of experiments performed with the AER 2-D model.

Substance Injected	Injected Particle Radius	Injection Rate (Tg yr ⁻¹)	Fractal Dimension D _f	Comments
Alumina	R ₀ = 80 nm ^a	1, 2, 4, 8	1.6, 2.8	
Alumina	R ₀ = 160 nm ^a	1, 2, 4, 8	1.6, 2.8	
Alumina	R ₀ = 240 nm ^a	1, 2, 4, 8, 16	1.6, 2.8	
Alumina	R ₀ = 320 nm ^a	1, 2, 4, 8	1.6, 2.8	
Diamond	R ₀ = 160 nm ^a	1, 2, 4, 8	1.6	
SO ₂	Gas phase	1, 2, 4, 10 ^c		Replication of Pierce et al. (2010)
H ₂ SO ₄	R _g = 95 nm, σ = 1.5 ^b	1, 3, 6, 15 ^c		Replication of Pierce et al. (2010)

2 ^a monomers of uniform radius

3 ^b lognormal distribution defined by mode radius, R_g, and width, σ, representing distribution after plume processing

4 ^c based on molecular weight of SO₂ or H₂SO₄, not S alone.

5

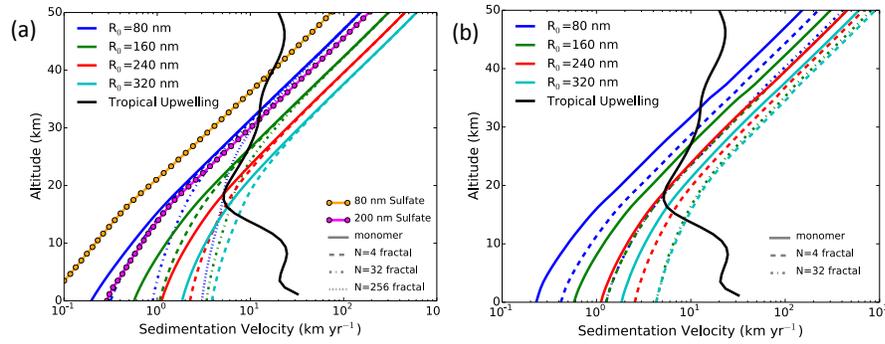
1 Table 2. Comparison of alumina, diamond, and sulfate solar geoengineering, based on a top-of-
 2 atmosphere shortwave radiative forcing of -2 W m^{-2} for each case.

Metric	Alumina 240 nm	Diamond 160 nm	Sulfate as H_2SO_4	Sulfate as SO_2	Comments
SW Radiative forcing per unit injected mass flux ($\text{W m}^{-2} (\text{Tg yr}^{-1})^{-1}$)	-0.26	-0.42	-0.25 ^a	-0.20 ^a	Other than diamond, the differences are minor.
Ozone impact (global average column change)	-5.6%	-6.1% to -7.1% ^b	-13% ^c	-11% ^c	Alumina and diamond have less ozone depletion than sulfates, though there is considerable uncertainty.
Diffuse light increase (W m^{-2})	10.1	6.3	21	19	Alumina and diamond are both better (less diffuse light) than sulfates. Exact results would require a more sophisticated radiative transfer model.
LW+SW Heating rate (K day^{-1}) in tropical lower stratosphere	0.052- 0.060 ^d	0.007- 0.010 ^d	0.22	0.30	Alumina is probably better (less heating) than sulfates, but this estimate is subject to considerable uncertainty. Diamond is much better.

3 ^a Sulfate emission fluxes based on mass of H_2SO_4 or SO_2 injected annually. [Sulfate fluxes needed to achieve a](#)
 4 [given RF would increase by approximately 20% if the LW RF component were included, decreasing the RF](#)
 5 [efficiencies shown here.](#)

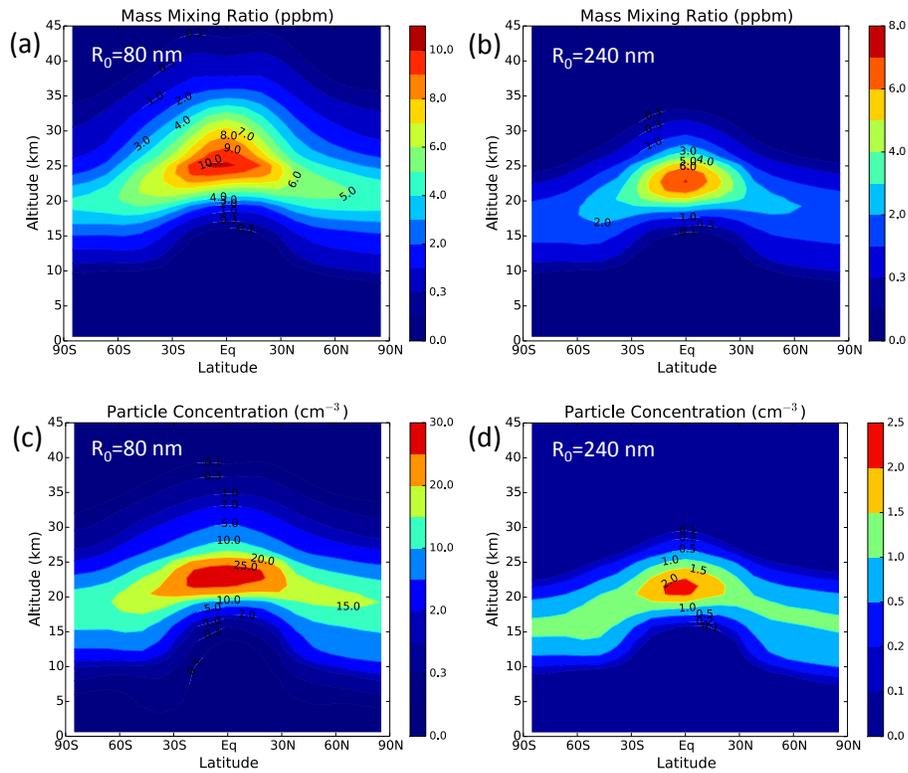
6 ^b The results for diamond are a range based on two cases, with and without reaction (R1) occurring on bare diamond
 7 surfaces.

- 1 ^c Note that the overall ozone loss from H₂SO₄ and SO₂ injection is higher than reported in most previous studies
- 2 because we consider short-lived bromine species
- 3 ^d The range for alumina and diamond is based on average particle number densities of monomers and of total
- 4 particles.



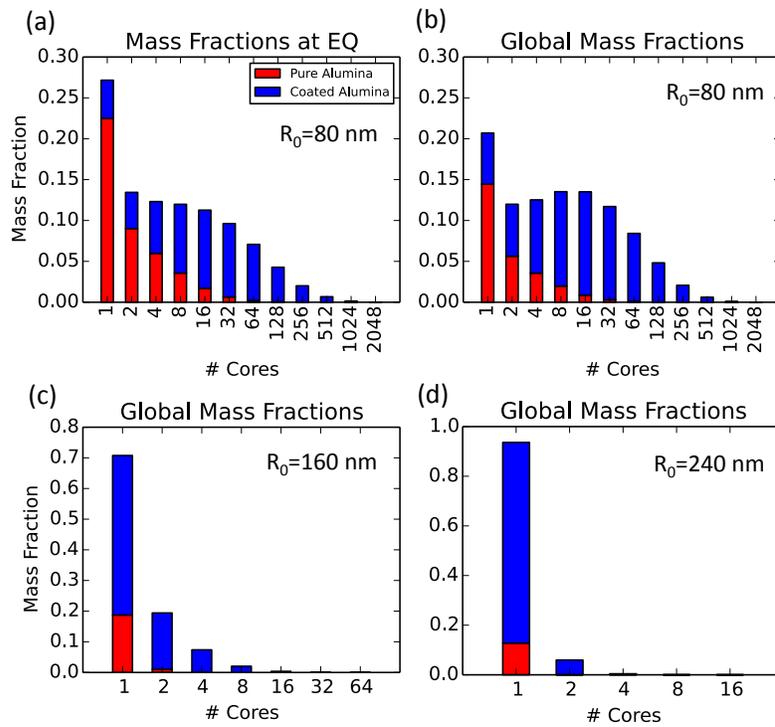
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2 Figure 1. Annual average sedimentation velocity (km yr⁻¹) versus altitude for (a) uncoated
 3 alumina particles and pure sulfate particles and (b) sulfate-coated aged alumina particles with
 4 compact fractal structures averaged over the region from 20°S to 20°N latitude. Solid colored
 5 lines represent monomers, dashed lines fractals with N=4, dash-dot lines fractals with N=32, and
 6 dotted lines fractals with N=256 (for R₀=80 nm only). Fractal dimension D_f=1.6 for uncoated
 7 particles represented in panel (a), D_f=2.8 for coated and compacted particles shown in panel (b).
 8 The black lines represent the annual average upwelling velocity of the model's advective
 9 transport averaged over the region from 20°S to 20°N latitude for comparison.



1

2 Figure 2. Mass mixing ratio of alumina in ppbm (panels a and b) and number density of alumina
 3 particles in cm^{-3} (panels c and d) with geoengineering injections of 1 Tg yr^{-1} of 80 nm monomers
 4 (panels a and c) and 1 Tg yr^{-1} of 240 nm monomers (panels b and d) for annual average
 5 conditions.

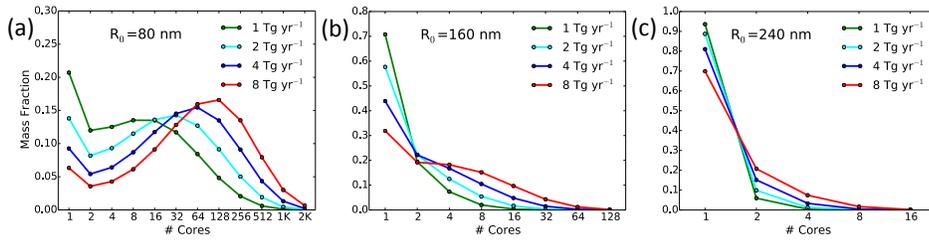


1

2 Figure 3. Distribution of integrated stratospheric alumina mass into monomers and fractals for
 3 geoengineering injection of 1 Tg yr^{-1} of alumina as 80 nm monomers at (a) the equator and (b)
 4 globally integrated, and for injection of 1 Tg yr^{-1} of alumina as (c) 160 nm and (d) 240 nm
 5 monomers globally integrated. Red bar length represents the mass fraction in dry alumina and
 6 blue bar length the mass fraction in coated alumina. Annual average conditions are represented.

7

1

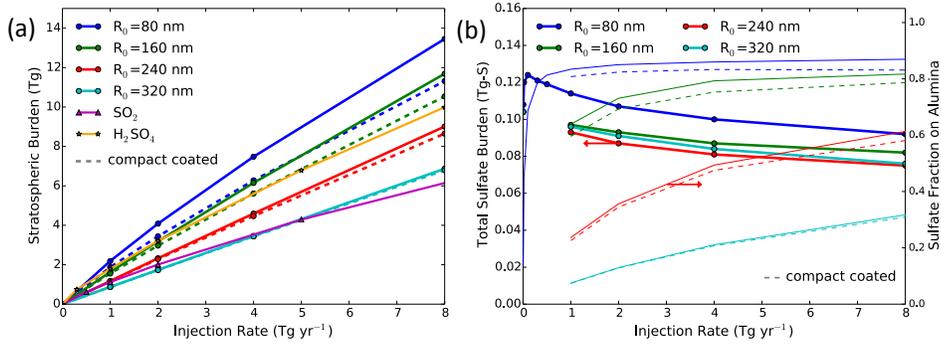


2

3 Figure 4. Calculated global annual average stratospheric mass fractions of alumina as a function
4 of the number of monomers contained in a fractal particle for (a) monomer injections of 80 nm
5 radius, (b) monomer injections of 160 nm, and (c) monomer injections of 240 nm radius, with
6 emission rates ranging from 1 to 8 Tg yr⁻¹.

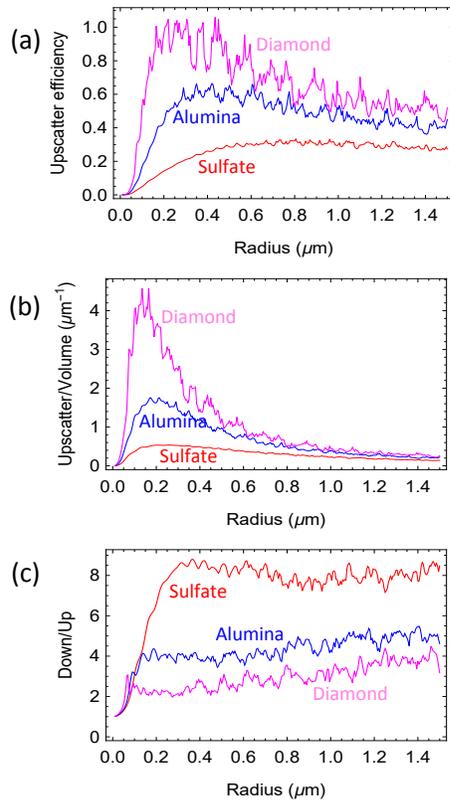
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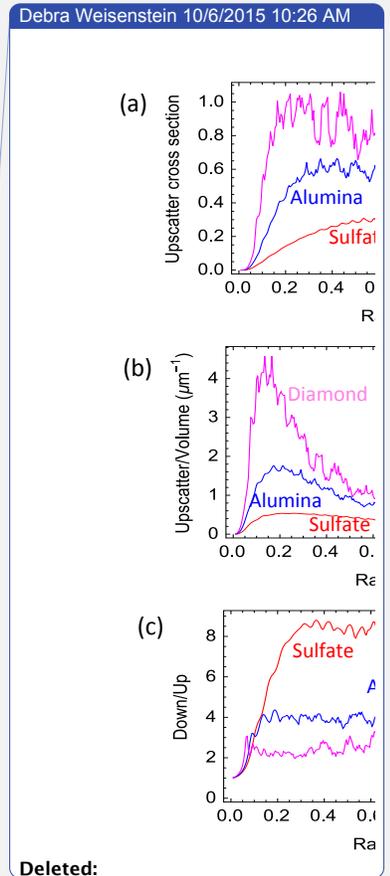
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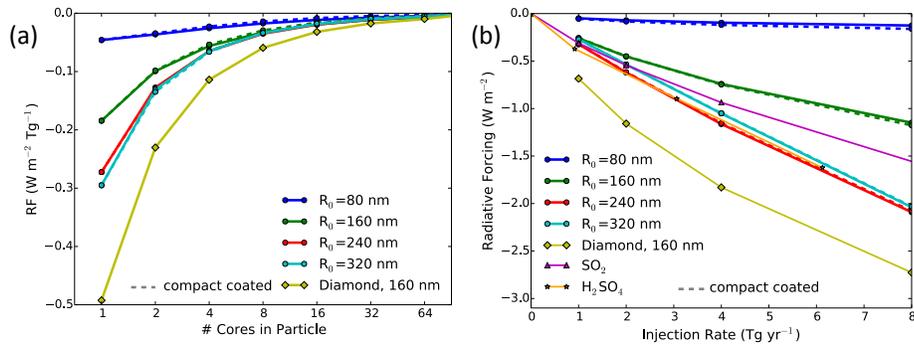
3 Figure 5. Annual average stratospheric burden of (a) alumina and (b) condensed sulfate versus
4 injection rate for various sizes of injected alumina monomers. For comparison, we plot sulfate
5 burden in Tg-S as a function of the rate of injection of SO₂ and H₂SO₄ (Pierce et al., 2010) in Tg-
6 S yr⁻¹ along with alumina burden in (a). Panel (b) shows the fate of natural sulfate as a function
7 of alumina injection rate, where the total sulfate burden is plotted on the left-hand axis (thick
8 lines with circles) and the fraction of that burden that is on the alumina particles is shown on the
9 right-hand axis (thin lines). The dashed lines represent simulations in which the coated alumina
10 particles are assumed to become more compact in shape.



1

2 Figure 6. Comparison of radiative scattering properties of alumina and diamond monomers and
 3 sulfate aerosol particles as functions of particle radius. Panel (a) shows the [upscatter efficiency](#)
 4 [which is the](#) upscatter cross-section divided by the geometric cross-section (a dimensionless
 5 ratio). Panel (b) shows the upscatter cross-section divided by the particle volume (units of μm^{-1}),
 6 and panel (c) shows the ratio of downscatter cross-section to upscatter cross-section integrated
 7 over the solar band.



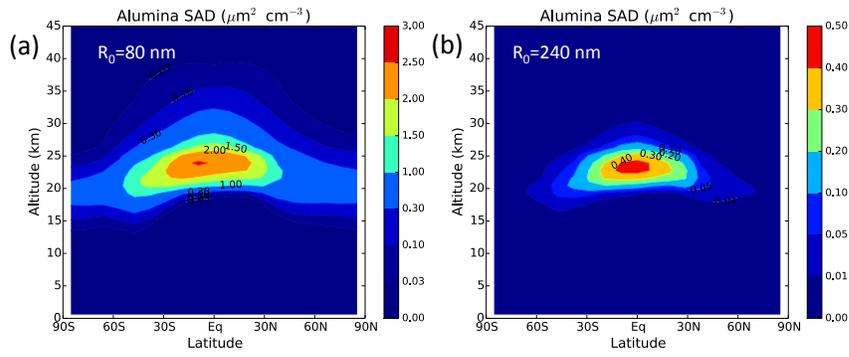


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2 Figure 7. Shortwave globally-averaged clear-sky radiative forcing per terragram burden (W m^{-2}
 3 Tg^{-1}) of alumina or diamond particles as a function of the number of monomer cores per fractal
 4 particle (panel a). Calculated globally-averaged shortwave radiative forcing as a function of
 5 injection rate for geoengineering scenarios (panel b) for annual average cloud-free conditions.
 6 The dashed lines represent simulations in which the coated alumina particles are assumed to
 7 become more compact in shape.

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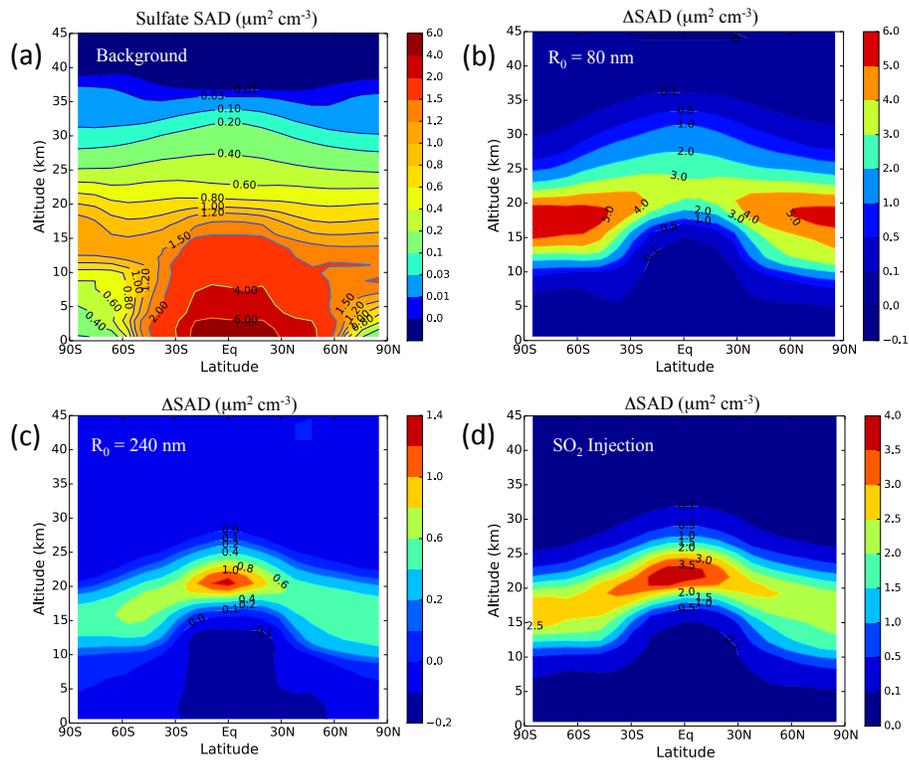
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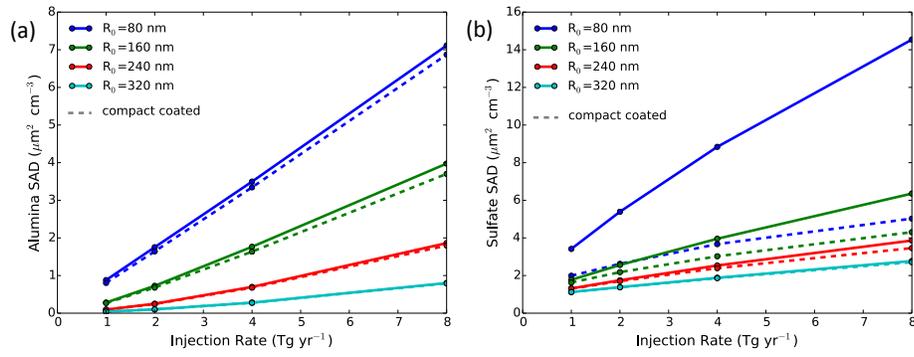
3 Figure 8. Calculated annual average surface area density ($\mu\text{m}^2 \text{cm}^{-3}$) of uncoated alumina
4 particles due to geoengineering with 1 Tg yr^{-1} injection of (a) 80 nm alumina monomers and (b)
5 240 nm alumina monomers.
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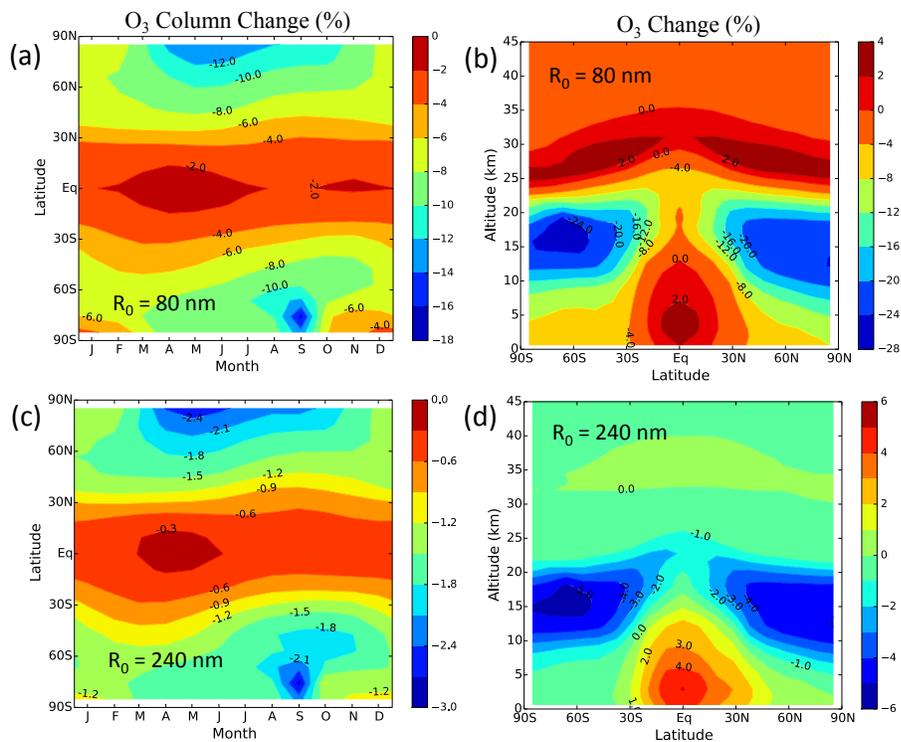
3 Figure 9. Calculated annual average sulfate surface area density ($\mu\text{m}^2 \text{cm}^{-3}$) of (a) sulfate
4 particles without geoengineering, and surface area density increase ($\mu\text{m}^2 \text{cm}^{-3}$) with
5 geoengineering injections of (b) 1 Tg yr^{-1} of 80 nm alumina monomers and (c) 1 Tg yr^{-1} of 240
6 nm alumina monomers. Panel (d) shows sulfate aerosol surface area density increase ($\mu\text{m}^2 \text{cm}^{-3}$)
7 with 1 Tg yr^{-1} of SO_2 injection.



1

2 Figure 10. Global annual average stratospheric surface area density between 15 and 25 km
 3 altitude for (a) uncoated alumina, and (b) total sulfate. The dashed lines represent simulations in
 4 which the coated alumina particles are assumed to become more compact in shape.

5

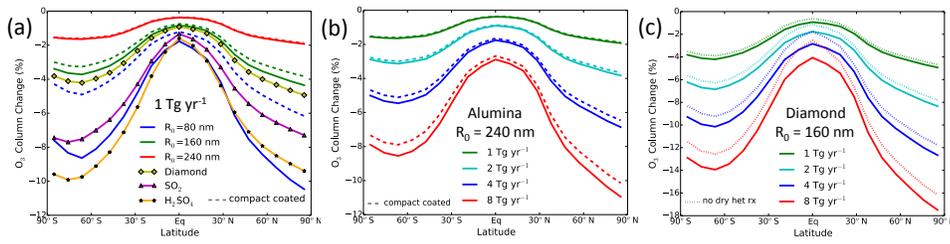


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2 Figure 11. Ozone changes due to injection of alumina aerosol. Column ozone changes (%) are
 3 shown as a function of latitude and month (left panels) and annual average local ozone changes
 4 (%) as a function of latitude and altitude (right panels). Results are shown for an injection rate of
 5 1 Tg yr^{-1} of 80 nm (top panels) and 240 nm (bottom panels) alumina monomers. Note ozone
 6 increases in the upper stratosphere where the NO_x cycle dominates and decreases in the lower
 7 stratosphere where the ClO_x and BrO_x cycles dominate.

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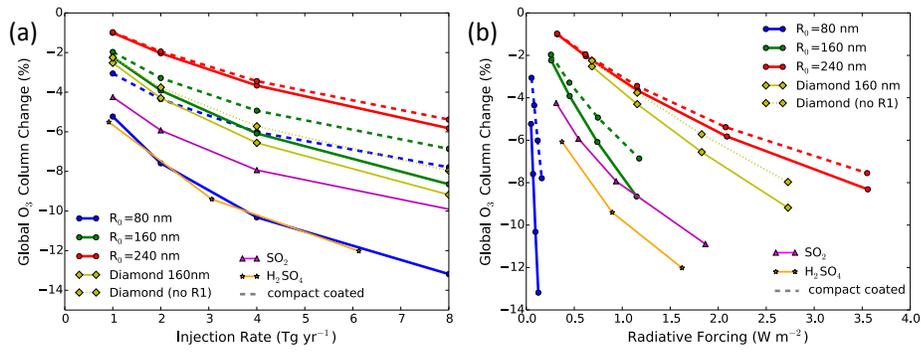


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3 Figure 12. Annual average column ozone change in percent as a function of latitude for (a) cases
4 with 1 Tg yr⁻¹ injections of alumina monomers of 80 nm, 160 nm, and 240 nm, and diamond
5 monomers of 160 nm and SO₂ and H₂SO₄, (b) cases with injection of 240 nm alumina monomers
6 at rates of 1, 2, 4, and 8 Tg yr⁻¹, and (c) cases with injection of 160 nm diamond monomers
7 at rates of 1, 2, 4, and 8 Tg yr⁻¹. Cases in which coated particles are assumed to become more
8 compact in shape are shown with dashed lines in panels (a) and (b). For diamond, cases without
9 reaction (R1) occurring on dry diamond particle surfaces are shown with dotted lines in panel
10 (c).

11

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3 Figure 13. Global annual average column ozone change (in percent) (a) as a function of
4 injection rate and (b) as a function of associated radiative forcing. Ozone change for diamond is
5 shown with and without reaction (R1) on uncoated diamond particles. Calculations with SO₂
6 and H₂SO₄ injections employ the same model to calculate radiative forcing and ozone depletion
7 as for alumina and diamond.