

[1,2]A. T.Lambe [2,\*]P. S.Chhabra [1,2]T. B.Onasch [3]W. H.Brune [4]J. F.Hunter [4]J. H.Kroll  
[1]M. J.Cummings [1]J. F.Brogan [1]Y.Parmar [2]D. R.Worsnop [2]C. E.Kolb [1]P.Davidovits  
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# ~~Comparison of secondary organic aerosol formed with an aerosol flow reactor and environmental reaction chambers: Effect of oxidant concentration, exposure time and seed particles on secondary organic aerosol chemical composition and yield~~

Chemistry Department, Boston College, Chestnut Hill, MA, USA

Aerodyne Research Inc., Billerica, MA, USA

Department of Meteorology and Atmospheric Sciences, The Pennsylvania State University, State College, PA, USA

Department of Civil and Environmental Engineering, Massachusetts Institute of Technology, Cambridge, MA, USA

now at: Department of Chemical Engineering, University of Texas, Austin, Texas, USA

*Correspondence to:* A. T. Lambe (lambe@aerodyne.com)

**Abstract.** We performed a systematic intercomparison study of the chemistry and yields of SOA generated from OH oxidation of a common set of gas-phase precursors in a Potential Aerosol Mass (PAM) continuous flow reactor and several environmental chambers. In the flow reactor, SOA precursors were oxidized using OH concentrations ranging from  $2.0 \times 10^8$  to  $2.2 \times 10^{10}$  molec  $\text{cm}^{-3}$  over exposure times of 100 s. In the environmental chambers, precursors were oxidized using OH concentrations ranging from  $2 \times 10^6$  to  $2 \times 10^7$  molec  $\text{cm}^{-3}$  over exposure times of several hours. The OH concentration in the chamber experiments is close to that found in the atmosphere, but the integrated OH exposure in the flow reactor can simulate atmospheric exposure times of multiple days compared to chamber exposure times of only a day or so. In most cases, for a specific SOA type the most-oxidized chamber SOA and the least-oxidized flow reactor SOA have similar mass spectra, oxygen-to-carbon and hydrogen-to-carbon ratios, and carbon oxidation states at integrated OH exposures between approximately  $1 \times 10^{11}$  and  $2 \times 10^{11}$  molec  $\text{cm}^{-3}$  s, or about 1-2 days of equivalent atmospheric oxidation. This observation suggests that in the range of available OH exposure overlap for the flow reactor and chambers, SOA elemental composition as measured by an aerosol mass spectrometer is similar whether the precursor is exposed to low OH concentrations over long exposure times or high OH concentrations over short exposure times. A linear correlation analysis of the mass spectra ( $m = 0.91\text{--}0.92$ ,  $r^2 = 0.93\text{--}0.94$ ) and carbon oxidation state ( $m = 1.1$ ,  $r^2 = 0.58$ ) of

20 ~~SOA produced in the flow reactor and environmental chambers for OH exposures of approximately~~  
 ~~$10^{11}$  molec cm<sup>-3</sup> s suggests that the composition of SOA produced in the flow reactor and chambers~~  
~~is the same within experimental accuracy as measured with an aerosol mass spectrometer.~~ This  
similarity in turn suggests that both in the flow reactor and in chambers, SOA chemical composition  
at low OH exposure is governed primarily by gas-phase OH oxidation of the precursors, rather than  
25 heterogeneous oxidation of the condensed particles. In general, SOA yields measured in the flow  
reactor are lower than measured in chambers for the range of equivalent OH exposures that can be  
measured in both the flow reactor and chambers. The influence of sulfate seed particles on isoprene  
SOA yield measurements was examined in the flow reactor. The studies show that seed particles  
increase the yield of SOA produced in flow reactors by a factor of 3 to 5 and may also account in  
30 part for higher SOA yields obtained in the chambers, where seed particles are routinely used.

## 1 Introduction

Laboratory and field studies over the last decade have shown that organic components of atmospheric  
particles constitute 20 to 50 % of the fine particle mass (PM) in the continental mid-latitudes, though  
the organic content can be higher (up to 90 %) in tropical forested regions (Kanakidou et al., 2005).  
35 On a global scale, 50–90 % of submicron organic PM is composed of oxygenated organic aerosol  
(Zhang et al., 2005) that is typically associated with secondary organic aerosol (SOA) formed by  
condensation of oxidized gas-phase species (Jimenez et al., 2009). Field studies indicate that SOA  
particles may influence cloud formation (Levin et al., 2014; Mei et al., 2013; Moore et al., 2012;  
Sihto et al., 2011) and may be optically active in the UV/Visible region of the electromagnetic  
40 spectrum (Zhang et al., 2011). These studies have also revealed the complexity of organic aerosol  
compositions and their chemical evolution via oxidative aging. The atmospheric lifetime of ambient  
SOA ranges from hours to weeks, providing a wide range of atmospheric exposures to a variety  
of oxidant species. Measured ambient SOA chemical compositions range from hydrocarbon-like  
organic aerosol, such as observed directly downwind of the Deepwater Horizon oil site during the  
45 2010 Gulf Oil Spill (Bahreini et al., 2012; De Gouw et al., 2011), to highly oxygenated OA, such as  
background SOA observed worldwide (Zhang et al., 2007). Much of this complexity is due to the  
thousands of organic compounds found in atmospheric particulate matter, specifically low volatility,  
highly functionalized species (Hallquist et al., 2009).

Laboratory experiments conducted in environmental chambers have been essential in providing  
50 SOA physical and chemical properties as well as yield data for predicting the rate of atmospheric  
SOA formation due to oxidation of biogenic and anthropogenic volatile organic compounds (VOCs)  
(Spracklen et al., 2011). Substantial progress has been made in understanding reaction mechanisms  
and the factors that influence SOA yields and composition. For example, SOA yields appear to have  
a complex dependence on VOC : NO<sub>x</sub> ratio (Loza et al., 2014; Ng et al., 2007), precursor concen-

55 tration/volatility (Presto and Donahue, 2006), and oxidant exposure (Lambe et al., 2012). Modeling  
observed atmospheric SOA levels therefore remains a challenge (Shrivastava et al., 2011; Bergström  
et al., 2012; Jo et al., 2013; Li et al., 2013) because of the large number of modeling parameters  
and associated sensitivities that are required to capture mechanistic details of SOA formation (Chen  
et al., 2013; Fountoukis et al., 2014).

60 Oxidant exposure is the integral of the oxidant species concentration and the sample residence  
time. Relatively low oxidant exposures are a major limitation of current environmental chamber  
techniques, which operate at OH concentrations ranging from approximately  $10^6$  to  $10^7$  molec $\text{cm}^{-3}$   
that are equal to or slightly more than daytime atmospheric OH concentrations. Chamber experi-  
ments are generally limited to residence times of several hours due to chamber deflation and the loss  
65 of oxidized vapors (Zhang et al., 2014) and/or particles to the chamber walls. This combination of  
low OH concentrations and residence time limits environmental chambers to simulating atmospheric  
aerosol particle lifetimes only up to 1 or 2 days, including the characterization of SOA yields. This  
limitation prevents the formation and the study of highly oxygenated SOA that is characteristic of  
aged atmospheric organic aerosol PM (Ng et al., 2010).

70 Recently, aerosol flow reactors have been developed to study SOA formation and evolution equiv-  
alent to multiple days of atmospheric OH exposure. In these reactors OH concentrations are typically  
 $\sim 10^9$  molec $\text{cm}^{-3}$  or greater, with reactor residence times of seconds to minutes (Kang et al., 2007;  
Hall IV et al., 2013; Keller and Burtscher, 2012; Lambe et al., 2011a; Slowik et al., 2012). With this  
range of OH concentrations and exposure times, flow reactors can simulate the full range of ambient  
75 levels of oxidation, measuring changes in SOA composition and yields over a wide range of equiva-  
lent atmospheric oxidation. Further, because of the short flow reactor residence times, experimental  
runs can be conducted on the scale of minutes rather than hours.

While flow reactors appropriately simulate the full range of integrated atmospheric oxidant expo-  
sures, in view of their short residence times and high oxidant concentrations it must be established  
80 how well the atmospheric aerosol chemistry is simulated (Renbaum and Smith, 2011). A growing set  
of studies indicates that flow reactor-generated SOA particles have compositions similar to ambient  
SOA, suggesting that the dominant oxidation reaction pathways in flow reactors are similar to those  
in ambient conditions (Bahreini et al., 2012; Lambe et al., 2012, 2011b; Kang et al., 2011; Massoli  
et al., 2010; Ortega et al., 2013; Slowik et al., 2012; Wang et al., 2012; Wong et al., 2011). [Other](#)  
85 [studies have used a combination of aerosol flow reactors and environmental chambers to character-  
ize heterogeneous uptake of organics on seed particles \(Jang et al., 2003\), SOA formation potential  
\(Kang et al., 2007; Bernard et al., 2012\), and evolution of functional groups in SOA with aging  
\(Ofner et al., 2010\); in general, similar results are obtained in reactors and chambers.](#) However,  
these comparisons need to be extended over a wider range of reactants and experimental conditions  
90 than are currently available.

Here we describe systematic intercomparison studies of SOA chemistry and yields generated from

a common set of precursors in a Potential Aerosol Mass (PAM) flow reactor (Lambe et al., 2011a) and four environmental chambers. SOA precursors studied are gas-phase alkane, biogenic, and aromatic compounds. SOA chemical composition and yield were characterized as a function of OH exposure. Additionally, the effect of sulfate seed particles on isoprene SOA yields was studied. Due to the limited oxidative exposure provided by the environmental chambers, direct comparison between the two techniques is possible only over a narrow range. However, reasonable extrapolations extend the range of interest.

## 2 Experimental

This manuscript compares properties of SOA produced in the PAM reactor to SOA produced in environmental chambers operated at the four institutions: California Institute of Technology (Caltech), Massachusetts Institute of Technology (MIT), Paul Scherrer Institut (PSI), and Carnegie Mellon University (CMU). The PAM reactor is a horizontal 13.3 L glass cylindrical chamber, 46 cm long  $\times$  22 cm ID and is operated in continuous flow mode with an average residence time of 100 s. The RH in the reactor was controlled in the range of 30–40 %. The Caltech, MIT, PSI and CMU Teflon chambers range from 7.5 to 28 m<sup>3</sup> in volume and are operated in batch or semi-batch mode with experimental residence times ranging from 4 to 10 h. The relative humidity (RH) in the Caltech, MIT and CMU chamber experiments was less than 10 %, and the RH in the PSI chamber experiments was controlled in the range of 40–50 %. A summary of the methods used for OH radical generation, particle generation, and data analysis is provided below.

### 2.1 OH radical generation

In the flow reactor, OH radicals were produced in the absence of NO<sub>x</sub> via the reaction  $O(^1D) + H_2O \rightarrow 2OH$ , with  $O(^1D)$  radicals produced from the reaction  $O_3 + h\nu \rightarrow O_2 + O(^1D)$ . O<sub>3</sub> (15–30 ppm) was generated by O<sub>2</sub> irradiation with a mercury lamp ( $\lambda = 185$  nm) outside the flow reactor. The  $O(^1D)$  atoms were produced by UV photolysis of O<sub>3</sub> inside the flow reactor using four mercury lamps which emit primarily at  $\lambda = 254$  nm. Additional photons are emitted at the following wavelengths with relative intensities of 1% or more of the UV intensity at 254 nm: 185 nm (1% Li et. al. (2015)); 302 nm (1%); 313 nm (1%); 366 nm (1%); 405 nm (1%); 436 nm (10%); 546 nm (1%) (BHK Inc. product specifications). At the highest UV intensity that was used in the reactor, we calculate upper-bound  $J_{UV} = 2 \times 10^{13}$  and  $2 \times 10^{15}$  cm<sup>-2</sup>s<sup>-1</sup> at  $\lambda = 185$  and 254 nm from ozone and OH exposure measurements. Corresponding lower limit timescales for UV photolysis of several phenols, carboxylic acids, aldehydes and ketones range from 12 to 50,000 sec for absorption cross sections ranging from approximately  $4 \times 10^{-17}$  to  $1 \times 10^{-20}$  cm<sup>3</sup> molec<sup>-1</sup> s<sup>-1</sup> (<https://sites.google.com/site/pamwiki/> and references therein).

In offline calibrations, OH concentrations were varied by changing the UV light intensity through

stepping the lamp voltages between 0 and 110V. SO<sub>2</sub> was added to the carrier gas, typically at mixing ratios ranging from 30 to 60 ppbv, and was used as an OH tracer. Calibrations were conducted at the same H<sub>2</sub>O and O<sub>3</sub> concentrations used in SOA experiments. At each lamp setting, OH exposures were quantified by measuring the steady-state SO<sub>2</sub> mixing ratio and normalizing to the SO<sub>2</sub> mixing ratio obtained with the lamps turned off. The corresponding OH exposure was quantified by measuring the decay of SO<sub>2</sub> normalizing the SO<sub>2</sub> mixing ratio with the lamps on to the SO<sub>2</sub> mixing ratio with the lamps off and applying the known OH + SO<sub>2</sub> rate constant (Davis et al., 1979), as shown in Equation 1:

$$\text{OH exposure} = -\frac{1}{k_{\text{SO}_2}^{\text{OH}}} \ln \left( \frac{[\text{SO}_2]}{[\text{SO}_2]_i} \right) \quad (1)$$

The concentrations ranged from approximately  $2.0 \times 10^8$  to  $2.2 \times 10^{10}$  molec cm<sup>-3</sup>. The corresponding OH exposures ranged from  $2.0 \times 10^{10}$  to  $2.2 \times 10^{12}$  molec cm<sup>-3</sup> s or approximately 0.2 to 17 days of equivalent atmospheric exposure. Additional SO<sub>2</sub> calibration measurements were conducted in the presence and absence of a subset of precursors (isoprene and JP-10) to investigate reductions in OH levels following addition of those precursors to the flow reactor at mixing ratios that were used in SOA experiments. No change in SO<sub>2</sub> decay was observed upon addition of isoprene, but addition of JP-10 decreased OH levels by approximately 10% (highest OH exposure) to 50% (lowest OH exposure) (Lambe et al., 2012). Reductions in OH exposure following addition of other VOCs will be investigated in future work using the methods of Li et. al. (2015).

In the environmental chambers, OH radicals were generated by UV photolysis ( $\lambda = 350$  nm) of hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) with no added NO<sub>x</sub>, or by UV photolysis of nitrous acid (HONO) or methyl nitrite (CH<sub>3</sub>ONO) with NO<sub>x</sub>. In the present studies, OH radicals generated in the Caltech chamber were formed from photolysis of H<sub>2</sub>O<sub>2</sub>, HONO, or CH<sub>3</sub>ONO, depending on the experiment, whereas OH radicals generated in the MIT, PSI and CMU chambers were formed exclusively from HONO photolysis. Typical chamber OH concentrations were approximately  $2 \times 10^6$  molec cm<sup>-3</sup> (H<sub>2</sub>O<sub>2</sub>) and  $2 \times 10^7$  molec cm<sup>-3</sup> (HONO) during the initial stage of chamber experiments. Corresponding OH exposures ranged from  $5.4 \times 10^{10}$  to  $4.0 \times 10^{11}$  molec cm<sup>-3</sup> s (Tab. 1), equivalent to approximately 0.4 to 3 days of atmospheric exposure at a typical 24 h average OH concentration of  $1.5 \times 10^6$  molec cm<sup>-3</sup> (Mao et al., 2009).

## 2.2 Particle generation

Figure ?? shows The gas-phase SOA precursors used in these studies include two biogenic compounds gases (isoprene,  $\alpha$ -pinene), three aromatic compounds (toluene, *m*-xylene, naphthalene), and three alkanes (n-C<sub>10</sub>, cyclodecane, tricyclo[5.2.1.0<sup>2,6</sup>]decane, also known as JP10). In the flow reactor, SOA was generated via gas-phase OH oxidation of precursors followed by homogeneous nucleation, or by condensation onto sulfuric acid or ammonium sulfate seed particles. The sulfuric

160 acid seed particles were generated by OH oxidation of SO<sub>2</sub> together with the SOA precursor, and ammonium sulfate seed particles were generated by atomizing an ammonium sulfate solution. The particles were dried and introduced continuously into the flow reactor ([without radioactive charge neutralization](#)) along with the gas-phase SOA precursor. In environmental chambers, SOA was generated via gas-phase OH oxidation of precursors usually followed by condensation onto ammonium sulfate seed particles. For long residence time chamber experiments, wall condensation of precursor gas-phase species can be significant. Seed particles are used in chamber studies to reduce wall effects. In some of the flow reactor experiments seed particles were also used to study their effect on SOA yields.

### 2.3 Particle monitoring and analysis

170 Particle number concentrations and size distributions were measured with a TSI scanning mobility particle sizer (SMPS). Aerosol mass spectra were measured with an Aerodyne time-of-flight aerosol mass spectrometers (ToF-AMS) (DeCarlo et al., 2006; Drewnick et al., 2005). Elemental analysis (Aiken et al., 2008) was performed on the AMS data to determine the bulk aerosol hydrogen-to-carbon (H/C) and oxygen-to-carbon (O/C) ratios along with the average aerosol carbon oxidation state ( $\overline{\text{OSc}}$ ) (Kroll et al., 2011). While AMS measurements provide basic information about SOA composition, additional supporting measurements [such as Fourier transform infrared spectroscopy, nuclear magnetic resonance, gas chromatography mass spectrometry, and chemical ionization mass spectrometry](#) are required to investigate SOA chemistry at the molecular level.

SOA yields were calculated from the ratio of aerosol mass formed to precursor gas reacted. The aerosol mass was calculated from the integrated particle volume and the effective particle density ( $\rho = D_{\text{va}}/D_{\text{m}}$ ), where  $D_{\text{va}}$  is the mean vacuum aerodynamic diameter obtained from the ToF-AMS and  $D_{\text{m}}$  is the electric mobility diameter obtained from the SMPS. Flow reactor SOA yields were corrected using size-dependent bis(2-ethylhexyl) sebacate wall-loss measurements (Lambe et al., 2011a); the average magnitude of these corrections was 32 % ( $\pm 15$  %) and represents an upper limit as it combines losses into and through the reactor. Caltech chamber yields were corrected for particle wall losses using size-dependent first-order loss coefficients determined from ammonium sulfate wall-loss measurements (Keywood et al., 2004). The magnitude of these particle wall loss corrections typically ranged from 10–30 %.

MIT chamber experiments were corrected for particle wall losses using the AMS organic-to-sulfate ratio to generate an upper limit and SMPS measurements of particle loss to generate a lower limit for aerosol yield (Hildebrandt et al., 2009). In the MIT chamber, these corrections were between a factor of 1.5 to 3.0 at the highest yields and OH exposures. Although the residence time in the flow reactor is much shorter than in the chambers, the surface-to-volume ratio in the PAM reactor is much greater. As a result particle losses are comparable in the two systems. Flow reactor SOA yields were also corrected for UV lamp-induced temperature increases by applying yield cor-

rections of  $-0.02$  per degree K of temperature rise (Qi et al., 2010; Stanier et al., 2007) relative to room temperature ( $\sim 293$  K). These temperature corrections ranged from 0–28 % (mean correction  $\pm 1\sigma = 7 \pm 7\%$ ). In the flow reactor, a known amount of precursor gas was introduced and the mass of reacted precursor gas was estimated from the OH exposure and known bimolecular rate constants (Atkinson, 1986). In environmental chamber studies, the mass of the remaining precursor gas was measured directly as a function of exposure time.

### 3 Results and discussion

#### 3.1 Sample mass spectra of flow reactor- and chamber-generated SOA

Figure 1 shows representative ToF-AMS spectra of SOA generated in the PAM reactor and the Caltech chamber (Chhabra et al., 2010, 2011) from the OH oxidation of  $\alpha$ -pinene and naphthalene, used here as representative biogenic and anthropogenic precursors respectively. The flow reactor spectra are obtained at an OH exposure of  $1.6 \times 10^{11}$  molec  $\text{cm}^{-3}\text{s}$ , or  $\sim 1$  day of equivalent atmospheric oxidation. The chamber spectra represent the SOA composition at peak aerosol formation ( $\sim 9 \times 10^{10}$  molec  $\text{cm}^{-3}\text{s}$  OH exposure). In this range the OH exposure for the PAM reactor and chamber are approximately the same, allowing for direct comparison. To quantify the similarity between mass spectra, we calculated the dot product between SOA mass spectra generated in the PAM flow reactor and the Caltech chamber (Murphy et al., 2003; Marcolli et al., 2006). Using this approach, each mass spectral signal is normalized to the square root of the sum of the squares of all signals in the mass spectrum. Each spectrum is represented as a normalized vector A or B, with dot product  $A \cdot B = \sum_{i=1}^n a_i b_i$ , where  $a_i$  and  $b_i$  are the normalized signals at each  $m/z$  in the spectrum;  $A \cdot B = 0$  indicates the spectra are orthogonal and  $A \cdot B = 1$  indicates the spectra are identical.

The top table inset in Fig. 1 shows the calculated dot products between each pair of mass spectra. The PAM flow reactor and the chamber produce particles with similar mass spectra, as indicated by dot products of 0.97 between spectra shown in Figs. 1a and c ( $\alpha$ -pinene SOA) and Figs. 1b and d (naphthalene SOA), suggesting similar compositions. The PAM flow reactor and the chamber produce particles with similar mass spectra, suggesting similar compositions. Features unique to  $\alpha$ -pinene and naphthalene SOA are observed in both flow reactor- and chamber-obtained spectra, with linear correlation coefficients of  $r^2 = 0.93$  ( $\alpha$ -pinene SOA) and  $r^2 = 0.94$  (naphthalene SOA) as noted in the bottom table inset in Fig. 1. For example,  $\alpha$ -pinene SOA spectra are dominated by signals at  $m/z = 43$  ( $\text{C}_2\text{H}_3\text{O}^+$ ) indicative of carbonyls, and several ion clusters below  $m/z < 100$  containing signals that are indicative of cycloalkyl fragments such as  $m/z = 27, 41, \text{ and } 55$ . On the other hand, AMS spectra of naphthalene SOA are dominated by  $m/z = 44$  ( $\text{CO}_2^+$ ), indicative of carboxylic acids, as well as signals that are indicative of aromatic compounds such as  $m/z = 50\text{--}51, 65, \text{ and } 76\text{--}77$ . As is evident from Fig. 1,  $\alpha$ -pinene and naphthalene SOA mass spectra display pronounced differences, with dot products ranging from 0.42 to 0.63 and  $r^2$  ranging from 0.18 to

0.37 between spectra shown in Figs. 1a and b, 1a and d, 1b and c, and 1c and d.

### 3.2 H/C, O/C ratios for flow reactor- and chamber-generated SOA

H/C and O/C ratios obtained from mass spectra such as shown in Fig. 2 provide information about the nature of SOA formation. Van Krevelen diagrams that show H/C ratios as a function of O/C ratios have been used to deduce oxidation reaction mechanisms for organic aerosols (Heald et al., 235 2010). Typically, with oxidative aging the O/C ratio increases and H/C ratio of SOA decreases as oxygen-containing functional groups are added to a carbon backbone. Here, we use Van Krevelen diagrams to compare the composition of SOA formed in the flow reactor and environmental chambers for the organic precursors studied. Direct comparisons are possible in the overlapping OH 240 exposure region. Typically the lowest OH exposures attained in the flow reactor overlap (or nearly overlap) with the highest OH exposures reached in environmental chambers (Table 1).

Figure 2 shows Van Krevelen diagrams obtained from laboratory SOA produced from the oxidation of gas-phase biogenic, aromatic, and alkane precursors. To simplify presentation, the data are displayed in three panels. Figure 2a shows biogenic SOA generated from isoprene and  $\alpha$ -pinene, 245 Fig. 2b shows SOA generated from aromatic compounds, and Fig. 2c shows SOA produced from alkanes. In most cases, for a specific SOA type the most-oxidized chamber SOA and the least-oxidized flow reactor SOA have similar Van Krevelen plots at integrated OH exposures between approximately  $1 \times 10^{11}$  and  $2 \times 10^{11}$  molec $\text{cm}^{-3}$ s, or about 1–2 days of equivalent atmospheric oxidation. This observation suggests that in the range of available OH exposure overlap for the flow 250 reactor and chambers, SOA elemental composition is similar whether the precursor is exposed to low OH concentrations over long exposure times or high OH concentrations over short exposures times. The flow reactor studies were done without added  $\text{NO}_x$ , whereas some of the environmental chamber studies were conducted in the presence of  $\text{NO}_x$ . The similarity in compositional parameters shown in Fig. 2 (e.g. H/C, O/C) were independent of the  $\text{NO}_x$  levels used in the environmental 255 chambers in the region studied, as has been observed in previous studies (Chhabra et al., 2011). The nitrogen-to-carbon (N/C) ratio ranged from 0.031 to 0.054 for SOA produced in the MIT chamber with added  $\text{NO}_x$  (Hunter et al., 2014) but was not characterized for other measurements shown in Fig. 2.

### 3.3 Carbon oxidation state for flow reactor- and chamber-generated SOA

260 Recently, the average carbon oxidation state ( $\overline{\text{OSc}}$ ), defined as  $\overline{\text{OSc}} = 2 \times \text{O/C} - \text{H/C}$ , was proposed as a more accurate indicator of atmospheric oxidative aging processes than the O/C ratio alone because this measure takes into account the level of saturation of the carbon atoms in the SOA (Canagaratna et al., 2014; Kroll et al., 2011). As will be demonstrated,  $\overline{\text{OSc}}$  of lightly oxidized SOA is strongly precursor-dependent. Figure 3 shows a scatter plot of  $\overline{\text{OSc}}$  for flow reactor and chamber 265 SOA for the eight gas-phase precursors studied. Different colored symbols are used to represent



each of the environmental chambers used in the intercomparison. For each data point obtained from environmental chamber measurements, we used data from the flow reactor obtained at the OH exposure that was closest in magnitude. A total linear least squares fit to the data presented in Fig. 3 ( $\text{PAM}_{\overline{\text{OSc}}} = 1.1 \cdot \text{Chamber}_{\overline{\text{OSc}}} - 0.16$ ;  $r^2 = 0.54$ ) indicates that there is no systematic  $\overline{\text{OSc}}$  difference observed across multiple SOA types produced in chambers and in flow reactors. For a specific SOA type, Figure 3 shows that the chambers and flow reactor provide similar  $\overline{\text{OSc}}$  for a specific SOA type with absolute differences ranging from 0.0040 to 0.60 (mean deviation =  $0.10 \pm 0.34$ ) over the range of measured SOA composition for comparable OH exposures. The observed deviations between PAM and chamber  $\overline{\text{OSc}}$  are no larger than deviations between two chambers (e.g.  $\alpha$ -pinene SOA produced in Caltech and PSI chambers, and cyclodecane SOA produced in MIT and CMU chambers).

### 3.4 SOA yields obtained in the flow reactor and environmental chambers

Several factors can affect SOA yields, including precursor concentration (Kang et al., 2011; Pfaffenberger et al., 2013; Presto and Donahue, 2006),  $\text{NO}_x$  (Presto et al., 2005; Ng et al., 2007), UV intensity/wavelength (Henry and Donahue, 2012), seed particle composition/loading (Ehn et al., 2014; Hamilton et al., 2011; Volkamer et al., 2009) and treatment of interactions between SOA and chamber walls (Hildebrandt et al., 2009; Matsunaga and Ziemann, 2010; Pierce et al., 2008; Zhang et al., 2014). For reference, the range of SOA precursor concentrations,  $\text{NO}_x$  levels, and seed particle concentrations used in SOA yield studies is summarized in Table 2. Isolation of individual factors is beyond the scope of this intercomparison. Because mass spectra and elemental ratios of SOA are similar whether it is generated in an environmental chamber or in a flow reactor (Sec. 3.1), we suggest that the differences in precursor concentration and UV wavelength (e.g.  $\lambda = 350$  nm vs.  $\lambda = 254$  nm) used in these studies have at most a minor effect on bulk composition.

Next we evaluate the influence of oxidant concentration and residence time on yields of SOA formed from common precursors in the PAM reactor and the Caltech and MIT environmental chambers. Figure 4 shows yields of SOA as a function of OH exposure for isoprene SOA (no added  $\text{NO}_x$ ),  $\alpha$ -pinene SOA (no added  $\text{NO}_x$ ), and tricyclo[5.2.1.0<sup>2,6</sup>]decane (JP-10) SOA, respectively. These precursors provide the broadest range of available yield values for intercomparison. Oxidation of isoprene forms SOA with low yield (Chhabra et al., 2010; Kroll et al., 2006), whereas  $\alpha$ -pinene forms SOA with moderate yields (Eddingsaas et al., 2012) and JP-10 forms SOA with mass yields greater than unity (Hunter et al., 2014; Lambe et al., 2012). Yields of alkane SOA do not display a systematic  $\text{NO}_x$  dependence (Loza et al., 2014); thus, to first order we assume different  $\text{NO}_x$  levels between the MIT chamber and PAM reactor do not influence our comparison of measured JP-10 SOA yields. The following features are evident in Fig. 4:

1. SOA yields at comparable OH exposures are a factor of 2 to 10 lower in the flow reactor than in chambers, whereas the mass spectra, O/C and H/C of SOA generated in the chambers and

flow reactor are similar (Figs. 1–3). Flow reactor SOA yields are also lower in the flow reactor than in chambers for the other precursors studied that are not shown in Fig. 4.

2. SOA yields in the flow reactor and chambers track each other; that is, the maximum yield of isoprene SOA is approximately 0.03 in the flow reactor and 0.06 in the Caltech chamber (Fig. 4a). Likewise, the maximum yield of JP-10 SOA is 1.4 in the flow reactor and 1.6 in the MIT chamber (Fig. 4c).
3. In the flow reactor, in all cases the SOA yield first increases as a function of OH exposure and then decreases. In some cases there is also evidence of a slight decrease in SOA yields at higher OH exposure in chambers (e.g. Fig. 4c).

One reason for the lower SOA yield in flow reactors may be the relative timescales for oxidation in the gas-phase vs. condensation onto pre-existing aerosols. The timescale for condensation of a gas-phase molecule onto pre-existing seed particles ( $\tau_{cond}$ ) can be calculated using Eq. 2 (Seinfeld and Pandis, 2006):

$$\tau_{cond} = \frac{1}{\alpha A_p} \sqrt{\frac{2\pi M_w}{k_B T}} \quad (2)$$

Where  $M_w$  is the molecular weight of the condensing species,  $\alpha$  is the mass accommodation coefficient, and  $A_p$  is the particle surface area. For example, over a representative range of particle surface area concentrations used in the flow reactor (10 to 100  $\mu\text{m}^2 \text{cm}^{-3}$ ), condensation timescales range from approximately 2000 to 20 000 s assuming a mass accommodation coefficient of 0.1 as has been measured for  $\alpha$ -pinene ozonolysis SOA (Saleh et al., 2013) and an average SOA molecular weight of 150  $\text{g mol}^{-1}$ . A lower limit of  $\tau_{cond} = 200$  to 2000 s is calculated over the same range of  $A_p$  assuming  $\alpha = 1$ . While our measurements do not constrain the mass accommodation coefficient, these timescales suggest that the residence time in the flow reactor (100 s) may not be adequate to allow complete condensation of semivolatile organic gas-phase species into SOA, whereas residence times in environmental chamber experiments are typically 10000 s or longer. Another factor in causing the SOA yield difference may be due to the condensation conditions. All the chamber experiments displayed in this work were done in the presence of ammonium sulfate seed particles, whereas seed particles were not normally used in our flow reactor studies. The effect of seed particles on SOA yields in the flow reactor is examined further in Sect. 3.5.

The observation that the yields track each other is a further indication that the reactive chemistry in the two systems is similar. The decrease in SOA yield subsequent to increase as a function of OH exposure is possibly due to gas-phase species carbon-carbon bond breaking from continued oxidation or heterogeneous OH oxidation reactions at high OH exposure (Hunter et al., 2014; Lambe et al., 2012; Loza et al., 2012); this trend is most clearly evident in the flow reactor studies. These

observations suggest that the first step in SOA formation is oxidation of gas-phase species leading to subsequent condensation. At low OH exposures equivalent to 1–2 days, heterogeneous reactions do not appear to play a significant role in SOA chemistry (Cappa and Wilson, 2012; Chen et al., 2013).

### 340 3.5 Effect of seed particles on SOA yields

The chamber experiments discussed in Sects. 3.1–3.4 were performed in the presence of ammonium sulfate seeds to promote more rapid condensation of gas-phase species into SOA. We investigated in more detail the influence of sulfate seed particle loading and composition on SOA yields produced in the flow reactor. We hypothesize that a “seed effect” should be most pronounced for SOA types  
345 with low yields, and investigate this proposal for isoprene SOA whereas the effect of sulfate seeds on  $\alpha$ -pinene SOA yields is minor (Kang et al., 2007). Figure 5 shows isoprene SOA yields in the flow reactor as a function of OH exposure using an ammonium sulfate seed particle mass concentration of  $20 \mu\text{g m}^{-3}$  which is comparable to sulfate volume concentrations used by Chhabra et al. (2010). At an OH exposure of  $2.9 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  or less in the flow reactor, the isoprene SOA yield with  
350 seeds added to the flow reactor compared to the SOA yield in the Caltech chamber is negligible. As the OH exposure is increased in the flow reactor, the SOA yield rises to a maximum value of 0.10 at  $7.8 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  OH exposure and then decreases. We note that the same trend was observed without adding seed particles, although with lower SOA yields. The O/C ratio of the flow reactor SOA decreases from 0.63 to 0.59 when seeds are added at  $7.8 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  OH exposure.  
355 Figure 5 supports the hypothesis that addition of seed particles promotes condensation, leading to higher SOA yields. Higher concentrations of seed particles may be required for condensation of gas-phase oxidation products to compete with continued OH oxidation in the gas phase.

To further investigate the effect of seeds on condensation, we measured isoprene SOA yields as a function of ammonium sulfate and sulfuric acid seed particle concentrations (Fig. 6). An OH  
360 exposure of  $7.8 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  was used in the flow reactor because this condition provided the best signal-to-noise ratio. It is evident from these figures that adding sulfate seeds significantly increases the SOA yield. At an OH exposure of  $7.8 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  and a sulfate seed particle concentrations of  $20 \mu\text{g m}^{-3}$ , the yield of isoprene SOA increases from 0.032 to approximately 0.14 in the presence of ammonium sulfate seeds and 0.25 in the presence of sulfuric acid seeds. **SMPS**  
365 **size distributions of the mixed particles suggest that most of the particle mass is measured by the AMS (Jayne et. a., 2000).** Increasing the seed particle concentration led to a continued increase in the yield, along with a decrease in the O/C ratio of the SOA as condensation of less-oxidized products was enhanced. The influence of seed acidity on isoprene SOA yields is well documented (Czoschke, 2003; Surratt et al., 2007; Offenberg et al., 2009), and the magnitude of the SOA yield enhancement  
370 in the presence of acidic seeds relative to neutral seeds in our work (factor of 2–3) is similar to these previous studies. Because a systematic isoprene SOA yield enhancement in the presence of neutral seeds (relative to unseeded conditions) is not observed in chamber studies (Chhabra et al.,

2010; Brégonzio-Rozier et al., 2014), these measurements suggest seed particles are required in flow reactor measurements of isoprene SOA (and potentially other types of SOA as well) in order to more closely simulate condensation conditions in environmental chambers.

#### 4 Conclusions

We performed a systematic intercomparison study of the chemistry and yields of SOA generated from OH oxidation of a common set of gas-phase precursors in several environmental chambers and a flow reactor. The most significant experimental parameters that varied between chambers and the flow reactor were OH concentration, residence time, and use of seed particles to promote condensation of oxidized vapors. The OH concentrations were 100 to 1000 times higher in the flow reactor, and residence times were 100 to 400 times higher in the environmental chambers. Within the range of approximate OH exposure overlap of  $(1-4) \times 10^{11}$  molec cm<sup>-3</sup> s, the SOA mass spectra and oxidation state were similar in both systems. The SOA yields for representative systems tracked each other but were lower in the flow reactor, probably in part because chamber SOA experiments were done with seed particles to promote condensation of oxidized vapors. Because SOA composition appears to be governed primarily by gas-phase OH oxidation processes, our results suggest that either flow reactors or chambers are properly characterizing SOA oxidative aging mechanisms representative of ambient conditions. However, SOA yields appear highly sensitive to the relative timescales of gas-phase OH oxidation and condensation processes. Simple calculations and measurements with seed particles suggest that condensation processes may be residence-time-limited in flow reactors depending on the mass accommodation coefficient of the oxidized vapors onto pre-existing particles. This could lead to underestimation of SOA yields. On the other hand, running an environmental chamber or flow reactor under conditions where partitioning is overestimated relative to atmospheric conditions will lead to a corresponding overestimation of SOA yields. Environmental chambers are commonly used to constrain SOA yields at low OH exposures, but flow reactors are needed to constrain SOA yields at higher OH exposures than 1–2 days of equivalent atmospheric oxidation, because environmental chamber SOA yield measurements appear to significantly underestimate atmospheric SOA formation rates when extrapolated over multiple days of equivalent atmospheric oxidation.

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**Table 1.** Summary of PAM reactor and environmental chamber OH exposure conditions.

table

Precursor	PAM	Caltech	MIT	PSI	CMU	Refs
isoprene	$1.6 \times 10^{11}$ – $1.0 \times 10^{12}$	$9.5 \times 10^{10}$	–	–	–	1–3, 8
$\alpha$ -pinene	$2.0 \times 10^{10}$ – $2.2 \times 10^{12}$	$5.4 \times 10^{10}$ – $1.3 \times 10^{11}$	–	$9.0 \times 10^{10}$ – $4.0 \times 10^{11}$	–	1–4, 7–8, 10
toluene	$1.6 \times 10^{11}$ – $2.1 \times 10^{12}$	$1.5 \times 10^{11}$	–	–	–	1–3, 8
<i>m</i> -xylene	$4.1 \times 10^{10}$ – $2.1 \times 10^{12}$	$7.7 \times 10^{10}$ – $1.6 \times 10^{11}$	–	–	–	1–3, 7–8
naphthalene	$1.6 \times 10^{11}$ – $2.1 \times 10^{12}$	$2.3 \times 10^{11}$	–	–	–	2–3, 8, 10
<i>n</i> -C <sub>10</sub>	$1.6 \times 10^{11}$ – $2.1 \times 10^{12}$	–	$2.6 \times 10^{11}$	–	–	1, 5, 9
cyclodecane	$1.6 \times 10^{11}$ – $1.9 \times 10^{12}$	–	$1.8 \times 10^{11}$	–	$2.2 \times 10^{10}$	1, 5–6
JP-10	$1.3 \times 10^{11}$ – $2.0 \times 10^{12}$	–	$3.3 \times 10^{11}$ – $5.8 \times 10^{11}$	–	–	5, 9–10

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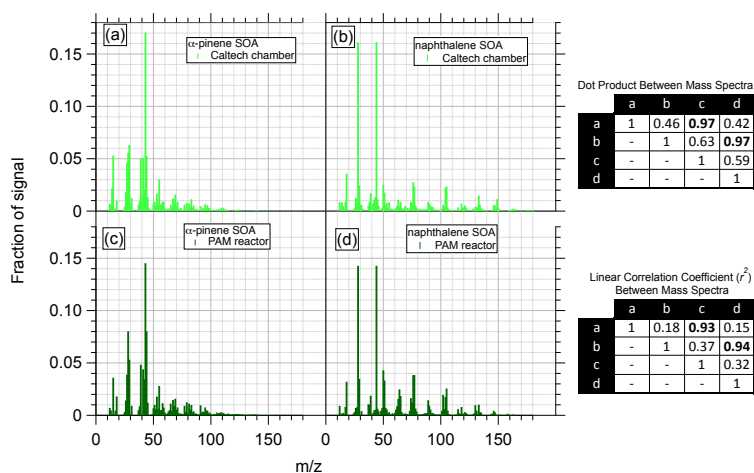
**Table 2.** Experimental conditions for PAM reactor, Caltech chamber, and MIT chamber yield measurements shown in Figs. 5–7.

	seed concentration ( $\mu\text{g m}^{-3}$ )	maximum [NO <sub>x</sub> ] added (ppb)	[isoprene] (ppb)	[ $\alpha$ -pinene] (ppb)	[JP-10] (ppb)	Refs
Caltech chamber	0–29 <sup>a</sup>	0	49–91	13.8–52.4	–	1–3
MIT chamber	50–100 <sup>a</sup>	475	–	–	42.9	4
PAM flow reactor	0–59 <sup>a</sup> ; 0–114 <sup>b</sup>	0	462	41–100	55	5–7

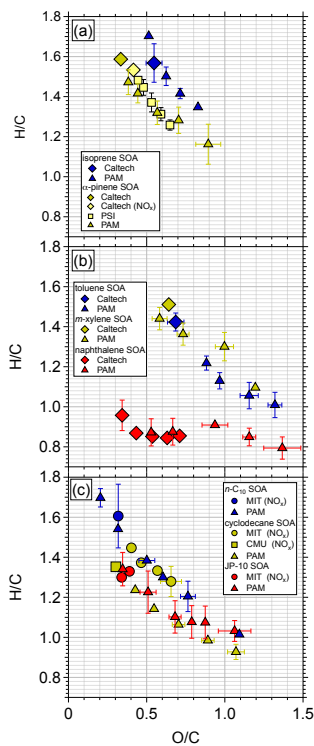
<sup>a</sup> ammonium sulfate seed;

<sup>b</sup> sulfuric acid seed;

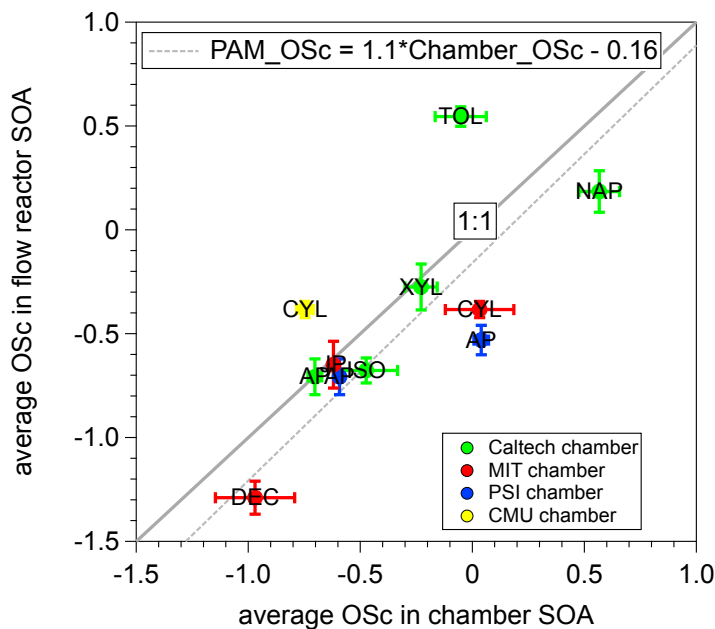
[1] Chhabra et al. (2010); [2] Ng et al. (2007); [3] Eddingsaas et al. (2012); [4] Hunter et al. (2014); [5] this work; [6] Lambe et al. (2012); [7] Chen et al. (2013).

**Fig. 1.** Aerodyne ToF-AMS spectra of SOA generated in the (a and b) Caltech environmental chamber and (c and d) PAM flow reactor from the OH oxidation of  $\alpha$ -pinene and naphthalene. Caltech chamber data obtained from Chhabra et al. (2011).

figure

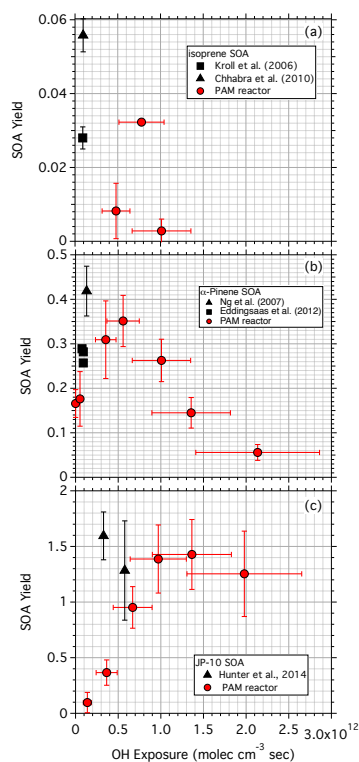


**Fig. 2.** Van Krevelen diagrams showing H/C ratio as a function of O/C ratio for SOA generated in the PAM flow reactor and environmental chambers by OH oxidation of (a) biogenic, (b) aromatic and (c) alkane precursors. Error bars indicate  $\pm 1\sigma$  uncertainty in binned O/C and H/C ratio measurements. Caltech, PSI, CMU, and MIT chamber data obtained from Chhabra et al. (2011); Pfaffenberger et al. (2013) (binned averages of O/C and H/C data from experiments # 1–9), Tkacik et al. (2012), and Hunter et al. (2014) respectively.

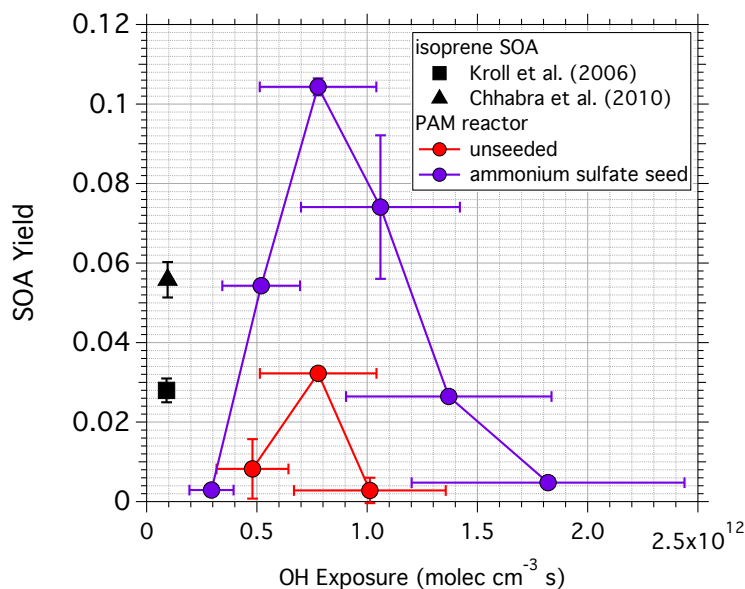


**Fig. 3.** Average carbon oxidation state ( $\overline{\text{OSc}}$ ;  $\overline{\text{OSc}} = 2 \times \text{O/C} - \text{H/C}$ ) of flow reactor- and environmental chamber-generated SOA. Error bars indicate  $\pm 1\sigma$  uncertainty in binned  $\overline{\text{OSc}}$  measurements. Markers indicate SOA precursor: TOL = toluene; NAP = naphthalene; XYL = *m*-xylene; CYL = cyclodecane; JP = JP-10; ISO = isoprene; AP =  $\alpha$ -pinene; DEC = *n*-C<sub>10</sub>. Caltech, PSI, CMU, and MIT chamber data obtained from Chhabra et al. (2011); Pfaffenberger et al. (2013); Tkacik et al. (2012), and Hunter et al. (2014) respectively.

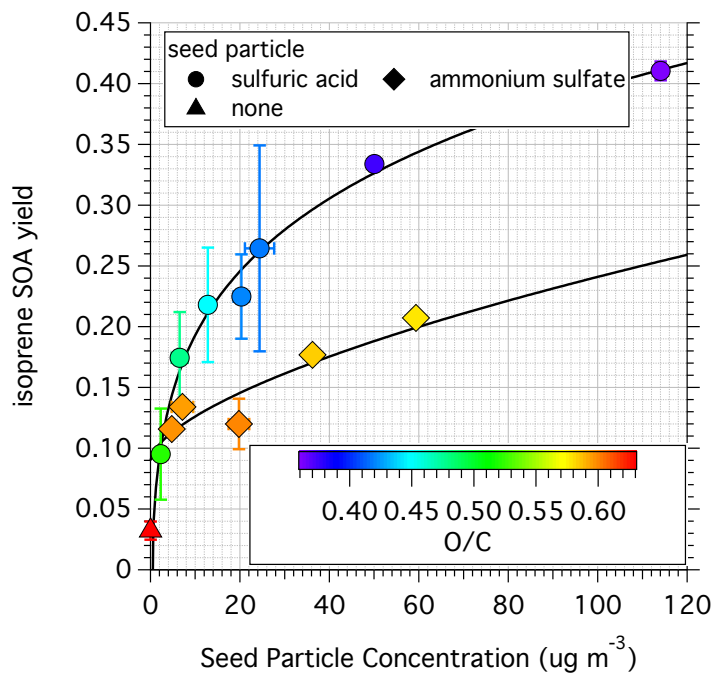




**Fig. 4.** Yields of SOA produced from photooxidation of (a) isoprene, (b)  $\alpha$ -pinene, and (c) tetracyclo[5.2.1.0<sup>2,6</sup>]decane (JP-10) in environmental chambers and PAM reactor as a function of OH exposure. The OH exposure in (c) is corrected for reductions in OH levels upon JP-10 addition (see Section 2.1). Error bars indicate  $\pm 1\sigma$  uncertainty in binned SOA yield measurements and 34% uncertainty in OH exposure values (Li et al., 2015). Black markers indicate data from Kröll et al. (2006); Chhabra et al. (2010); Eddingsaas et al. (2012), and Ng et al. (2007) obtained in the Caltech chamber and data from Hunter et al. (2014) was obtained in the MIT chamber.



**Fig. 5.** Yields of SOA produced from photooxidation of isoprene in the PAM reactor as a function of OH exposure in the presence of  $20 \mu\text{g m}^{-3}$  ammonium sulfate seed. Error bars indicate  $\pm 1\sigma$  uncertainty in binned SOA yield measurements and 34% uncertainty in OH exposure values (Li et. al., 2015).



**Fig. 6.** Yields of isoprene SOA produced in the PAM reactor at an OH exposure of  $7.8 \times 10^{11} \text{ molec cm}^{-3} \text{ s}$  as a function of seed particle concentration using ammonium sulfate and sulfuric acid seeds. Error bars indicate  $\pm 1\sigma$  uncertainty in binned measurements. Lines are power law fits to guide the eye.