We are very grateful that reviewer 1 was willing to review the manuscript a third time. Below you can find the review comments and our answers to them.

”In Section 3.1, Page 6, lines 15-20, they say that they represented only the initial reactions of OH, O3, and NO3 with some compounds whose subsequent reactions were not represented in MCM "considered as a sink of OH, O3, and NO3 without any other influence on ... gas-phase chemistry. However, at least the OH reaction are not really sinks for OH, they regenerate most of the OH back after some NO to NO2 conversions. The NO to NO2 conversions (which affect ozone formation and, indirectly, radical levels) are also ignored in the O3 and NO3 reactions that are neglected. This isn't important in this application because the error in ignoring the regenerated OH doesn't matter, since they are using OH calculated from consumption rates in the model, and their main interest is calculation of rates of formation of particle precursors, not O3. However, it would have been less of an approximation to use reactions of the most similar MCM species, as they did with the other compounds considered more important. They don't need to re-do the calculation, but just point out that this approach affects accuracy in the model in predicting radicals and ozone. Note that the discrepancies shown on Figure 3 cannot be explained by this omission in the model.”

On Page 6 after line 15-22 we have added:

“To only consider the first oxidation reactions of some of the BVOCs influence the predicted radical and ozone concentrations. However, because the concentrations of these BVOCs were relatively low this effect was most likely small.”

“It might be useful to include in Table S1 a footnote indicating how the subsequent reactions were represented, which could be (1) based on MCM for this compound, (2) based on MCM for another compound (state which), or (3) only the initial reaction is represented, not the reactions of radicals or species formed. This is stated in the text, but having the footnotes here is a good reference and may allow shortening the text if desired by the editor.”

Yes, thank you this is a good suggestion. We have added such footnotes to Table S1 but we have not removed any of the text in the manuscript.

“The curves with the yellow dashed lines on Figure 4, which seem to be exactly on top of the solid black line, should included in the legend on the figure as well as the figure caption.”

It is already included in the legend but it is hard to see the yellow color so we have change the color of the dashed line to orange.

“I didn't notice this in my previous review, but it looks like Table 2 has three columns. It might be helpful to have column headers in this case.”

Yes we agree, we have added the column headers:

Sensitivity test category	**Varied parameters (method) 1**	**Varied parameters (method) 2**

“In Section 4.2.1, page 22, lines 4-6 they state that their model simulations shown on Figure 6a with FEP parameters was about 2-3 times higher than observed. They should say this is for Day 1. For Days 2-4 the overprediction wasn't near as much (about 50% higher).”

Yes you are partly correct. We have made a more carful check of the difference between the model and observations for each day. With the FEP wall parameterization the model was about 3-4 times higher for Day-1, 1.5-2 times higher during Day 2 and 3 and 2-3 times higher for Day-4. In the manuscript we now write:

“Because of the lower VOC wall losses, the model overestimated the SOA formation by a factor of 3-4 for Day-1, a factor of 1.5-2 for Day-2 and Day-3 and a factor of 2-3 for Day-4.”

“It would have been helpful to me if Table 3 showed which simulations used "R1" and "R2" so the reader doesn't have to refer back to elsewhere in the text to see what this means. Either that use terms that correspond to the captions on Table 3 when discussing Table 3 and Figure 11 in the text. My understanding is that R1 corresponds to "SQT ox by O3" and R2 corresponds to "SQT ox by OH". If I am not understanding this correctly, then the discussion has to be clarified.”

R1 and R2 was used for the ELVOCs that contributes to the SOA formation. As a third sentence of Sect. 4.3.2 we write: “For all simulations in this section, ELVOCs were formed from all the monoterpenes and sesquiterpenes oxidized by O_3 and OH (R1-R2).”

R3 and R4 were used to represent the formation of $ELVOC_{nucl}$ that contributes to the nano-CN formation. We have added to Table 3 which reactions R3 and/or R4 that was used to represent the $ELVOC_{nucl}$ formation from monoterpens and sesquiterpenes. R3 correspond to "SQT ox by O3" and R4 to "SQT ox by OH".

The revised manuscript with tracked changes are found below:

Modelling the contribution of biogenic volatile organic compounds to new particle formation in the Jülich plant atmosphere chamber

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Abstract

We used the Aerosol Dynamics gas- and particle-phase chemistry model for laboratory CHAMber studies (ADCHAM) to simulate the contribution of BVOC plant emissions to the observed new particle formation during photooxidation experiments performed in the Jülich Plant-Atmosphere Chamber and to evaluate how well smog chamber experiments can mimic the atmospheric conditions during new particle formation events. ADCHAM couples the detailed gas-phase chemistry from Master Chemical Mechanism with a novel aerosol dynamics and particle phase chemistry module. Our model simulations reveal that the observed particle growth either may have been controlled by the formation rate of semi- and low-volatility organic compounds in the gas-phase or by acid catalyzed heterogeneous reactions between semi-volatility organic compounds in the particle surface layer (e.g. peroxyhemiacetal dimer formation). The contribution of extremely low-volatility organic gas-phase compounds to the particle formation and growth was suppressed because of their rapid

and irreversible wall losses, which decreased their contribution to the nano-CN formation and growth compared to the atmospheric situation. The best agreement between the modelled and measured total particle number concentration (R^2 >0.95) was achieved if the nano-CN was formed by kinetic nucleation involving both sulphuric acid and organic compounds formed from OH oxidation of BVOCs.

1 Introduction

New particle formation, including formation of nano condensation nucleii (nano-CN) (McMurry et al., 2011) and their growth to larger sizes, has been observed world-widely in continental boundary layers and free troposphere (Kulmala et al., 2004; Mirme et al., 2010). Field observations, laboratory experiments and model simulations indicate that gaseous sulphuric acid (H_2SO_4) plays an important role in atmospheric nano-CN formation, yet H_2SO_4 alone appears not to be able to explain all the steps of this process (Kulmala et al., 2000; Boy et al., 2003; Sipilä et al., 2010; Riipinen et al., 2007; Sihto et al., 2006; Kerminen et al., 2010; Kulmala et al., 2013, 2014). Basic compounds like ammonia and certain amines have been proposed to act as stabilizing compounds in nano-CN clusters (Berndt et al., 2010; Almeida et al., 2013; Kurtén et al., 2008), while subsequent steps of atmospheric new particle formation seem to rely on the presence of low-volatility organic compounds (LVOCs) (e.g., Metzger et al., 2010; Paasonen et al., 2010; Riipinen et al., 2012; Ehn et al., 2014; Schobesberger et al., 2013).

Oxidation products of biogenic volatile organic compounds (BVOCs) constitute the largest source of secondary organic aerosol (SOA) in the global atmosphere (Tsigaridis and Kanakidou, 2003; Hallquist et al., 2009; Spracklen et al., 2011), accounting for the main composition of SOA condensational growth (VanReken et al., 2006; Hao et al., 2009; Riipinen et al., 2012). BVOC oxidation also produces extremely low-volatility organic compounds (ELVOCs) essential to the new particle formation process in the atmosphere (Ehn et al., 2014). The most abundant group of BVOCs, accounting for more than half of their global emissions, are terpenoids (Guenther et al., 1995). Terpenoids include compounds consisting of one to several isoprene units, e.g. isoprene (C_5H_8), monoterpenes ($C_{10}H_{16}$), and sesquiterpenes ($C_{15}H_{24}$). Oxidation products of monoterpenes have substantial contribution to SOA formation (Hoffmann et al., 1997, 1998; Laaksonen et al., 2008), and low-volatility substances produced by sesquiterpene-ozone reactions may also initiate SOA formation (Bonn and Moortgat, 2003). Oxidation of isoprene leads to the formation of SOA (Surratt et al., 2006; Claeys et al., 2004), yet isoprene may also suppress the new particle formation process due to its high reactivity with OH (Kiendler-Scharr et al., 2009). Overall, the exact contribution of BVOCs to SOA still remains uncertain, especially with respect to the initial steps of atmospheric new particle formation.

In this study, we used the Aerosol Dynamics, gas- and particle-phase chemistry model for laboratory CHAMber studies (ADCHAM) (Roldin et al., 2014), to investigate the nano-CN formation and growth during a measurement campaign conducted in the Jülich Plant Atmosphere Chamber (JPAC). The in-depth analysis of the chamber measurements is discussed in another paper (Dal Maso et al., 2014). Here, we use the full chamber dataset including gas and particle phase measurements, either as model input or for evaluation of the model results.

The main objectives of this work were to evaluate how well the JPAC experiments could mimic the real atmospheric conditions during new particle formation events over the boreal forest and to constrain the dominating mechanisms responsible for the nano-CN formation and growth.

2 Measurement set up

The experiments were conducted in the JPAC located at Forschungszentrum Jülich, Germany. Detailed description regarding the chamber facility was given in previous articles (e.g., Mentel et al., 2009; Schimang et al., 2006). In addition, more details about this measurement campaign can be found from Dal Maso et al. (2014). In brief, the system consisted of two borosilicate glass chambers with PTFE Teflon floors. The chambers were operated as continuously stirred tank reactors (CSTR) with Teflon fans ensuring homogeneous air mixing. Each chamber housed adjustable temperature between 10 and 50 °C with a stability of ±0.5 °C. The small chamber (1150 L) served as plant chamber and was connected to the larger chamber that worked as reaction chamber (1450 L, surface-area-to-volume ratio 4.87 m^{-1}).

Three small trees aging from three to four years were brought from Hyytiälä, Finland, and included Scots pine (*Pinus sylvestris*), Norway spruce (*Picea abies*), and Silver birch (*Betula* *pendula*). These trees, representing the main boreal forest species in Finland, were placed in the JPAC plant chamber for almost two weeks before the intensive experiment campaign started. This allowed them to adjust to the chamber environment. Starting from the day when the trees were installed in the plant chamber, tree emissions were transferred into the reaction chamber, where O_3 was added together with water vapour directly and OH was generated periodically by turning on and off the UV light, in the same fashion as applied in the intensive phase. New particles were formed every day during the UV-light on periods and gases and particles deposited on the chamber walls. Discharge lamps (Osram HQI 400 W/D) were used for illumination to simulate the solar light spectrum in both chambers. Filters (OptoChem, type IR3) that reflect wavelengths between 750 and 1050 nm were used as heat shields to avoid infrared radiation inflicted plant overheating. For the model simulations we used the measured spectrum for the discharge lamps in the wavelength range of 280-650 nm and a single UV-light peak at a wavelength of 254 nm which represents the UV-spectrum from UV-light source, a Philips, TUV 40W lamp (Fig. S1). The UV-light source intensity corresponds to a O_3 to $O(^1D)$ photolysis rate of 2.9×10^{-3} s^{-1} (Mentel et al., 2009).

The ambient air was purified by an adsorption dryer (Zander, KEA 70) and a palladium catalyst (450 °C). O_3, NO, NO_2 and VOC levels decreased significantly after passing the purification system. The flow through the plant chamber was 115 L min^{-1} from which a fraction of 20 L min^{-1} was transferred to the reaction chamber. This flow was kept nearly constant by keeping the pressure drop constant. In addition, the 10 L min^{-1} flow containing ozone was added, controlled by a second flow controller.

The concentrations of O_3, CO_2 and H_2O were measured by commercial analytical instruments. Two Gas Chromatography Mass Spectrometer (GC-MS) systems were used, one to measure the VOC concentrations from C_5 to C_{20} in the outflow air from the plant chamber (Heiden et al., 2003), and another to identify the OH concentration by determining the decrease in the concentration of 2-butanol in the reaction chamber (Kiendler-Scharr et al., 2009). Meanwhile, the VOC concentration was continuously measured by an on-line Proton Transfer Reaction Mass Spectrometer (PTR-MS) in the plant and reaction chamber. The gas phase H_2SO_4 concentration in the reaction chamber was measured by a Chemical Ionization Mass Spectrometer (CIMS) (Petäjä et al., 2009; Mauldin et al., 1998). A prototype Airmodus Particle Size Magnifier (PSM) coupled with a TSI condensation particle counter (CPC) was used to count the total number concentration of particles larger than ca. 1.6 nm in diameter

(Vanhanen et al., 2011) and a TSI CPC (TSI3022A) was used to measure the total concentration of particles larger than ca. 7 nm in diameter. A Scanning Mobility Particle Sizer (SMPS TSI3071 + TSI3025A) was used to measure the particle size distribution in the size range of 14 – 600 nm.

The real plant emissions and the simulated day and night conditions make these experiments suitable for evaluation of methods used to describe the atmospheric transformation (ageing) of BVOCs and SOA beyond the first gas-phase VOC oxidation stage (e.g. the 2-Dimensional Volatility Basis Set (2D-VBS) Donahue et al., 2011).

3 Description of the ADCHAM model and its application

ADCHAM is a model primarily developed for simulations of laboratory chamber experiments on SOA formation and ageing. The model includes modules for reversible partitioning of organic compounds to and from the chamber walls, all fundamental aerosol dynamics processes, detailed gas- and particle-phase chemistry and a kinetic multilayer model which can be used to simulate mass transfer limited mixing of compounds in the particle phase (Fig. 1). Below we describe how ADCHAM was set up in this work. For a more detailed description of ADCHAM we refer to Roldin et al. (2014).

3.1 Gas-phase chemistry

The gas-phase chemistry reactions were selected from the Master Chemical Mechanism (MCM) v3.2 (Jenkin et al., 1997, 2012; Saunders et al., 2003) via website: http://mcm.leeds.ac.uk/MCM/. The MCM is a near-explicit chemical degradation mechanism that simplifies the chemical path of compounds by lumping products beyond the second oxidation step. We used the Kinetic Pre-Processor (KPP) version 2.1 (Damian et al., 2002) to simulate the gas-phase chemistry.

Among all the compounds measured by GC-MS in the JPAC plant chamber, 28 organic compounds were included in the gas-phase chemistry mechanism (see Table 1). In the table, the “other MTs” equals to the concentration difference between the summation of the concentrations of the 13 selected monoterpenes by GC-MS and the total monoterpene concentration measured by PTR-MS, while the “other SQTs” equals to the summation of other sesquiterpene isomers besides the listed four sesquiterpenes measured by GC-MS.

In total, the gas-phase chemistry mechanism has 2294 species and 6487 chemical and photochemical reactions for the chemical calculations. These include relevant inorganic reactions and the full MCM chemistry path for isoprene, *α*-pinene, *β*-pinene, limonene, *β*-caryophyllene, toluene, 2-butanol and hexanal. We also included the initial reactions of OH, O_3, and NO_3 with the following organic compounds: myrcene, sabinene, camphene, ocimene, Δ^3-carene, "other MTs" (which we assumed to have the same rate coefficients as *α*-pinene), cineole (also known as eucalyptol), farnesene, "other SQTs" (which assumed the same rate coefficients as *β*-caryophyllene), *α*-terpinene, Δ-terpinene, *α*-phellandrene, *β*-phellandrene, and terpinolene. Furthermore, we included the initial reactions between OH and the following organic compounds: tricyclene, nonanal, bornyl acetate and methyl salicylate. Finally, the initial reactions between O_3 and the following organic compounds were included: *α*-humulene, *α*-longipinene, and Δ-cardinene. Table S1 lists the reaction rate coefficients that were used for all 28 VOCs, together with references to the publications where the rate coefficients can be found. For the BVOCs with very small contribution to the total BVOC concentration in the chamber (myrcene, sabinene, camphene, cineole, *α*-terpinene, Δ-terpinene, *α*-phellandrene, *β*-phellandrene, terpinolene, tricyclene, nonanal, bornyl acetate, methyl salicylate*,* *α*-longipinene and *α*-humulene), the initial reaction was only considered as a sink of OH, O_3 and NO_3, without any other influence on the MCM gas-phase chemistry scheme. The oxidation of the sesquiterpenes farnesene, Δ-cardinene and "other SQTs" were treated in a similar manner, except that we assumed that on average ½ OH molecule was formed for each sesquiterpene molecule that reacted with O_3. To only consider the first oxidation reactions of some of the BVOCs influence the predicted radical and ozone concentrations. However, because the concentrations of these BVOCs were relatively low this effect was most likely small.

Ocimene constituted a major fraction of MT on the first experiment day. The chemistry path for ocimene is not available in the MCM. However, we included the ocimene chemistry by approximating its chemistry beyond the first oxidation step with that of limonene, for which the full MCM chemistry path is available. We also approximated the chemical path for Δ^3-carene after the first oxidation step with that of *α*-pinene, since both are bicyclic monoterpenes with an endocyclic double bond, as was done by Boy et al. (2013). For the "other MTs", we approximated the chemistry beyond the first oxidation step assuming that 50 % of the oxidation products end up as the corresponding MCM *α*-pinene first generation oxidation products and 50 % end up as the corresponding MCM *β*-pinene first generation oxidation products.”

Based on the recent finding of rapid formation of extremely low-volatility organic compounds (ELVOCs) from ozonolysis of monoterpenes containing endocyclic double bonds (Ehn et al., 2014) we also included a simplified ELVOC formation mechanism in the MCM gas-phase chemistry code (R1), assuming that 7 mole % of the α-pinene + O_3 and Δ^3-carene + O_3 oxidation products were ELVOCs, with a molar mass of 325 g/mol and a vapour pressure of 10^{-10} Pa, which approximately corresponds to the VBS bin of $\log_{10}(C^*/\ \mu g\ m^{-3}) = -5$. We also performed simulations where we considered that ELVOCs were formed from ozonolysis of any monoterpene or sesquiterpene, with the same ELVOC molar yield as for α-pinene. $MCM_{ox.prod}$ represents the MCMv3.2 oxidation products that were formed from the same reactions as the ELVOCs but via a different reaction pathway.

$$\mathrm{VOC}+\mathrm{O_3}\rightarrow 0.07\mathrm{ELVOC}+0.93\mathrm{MCM_{ox.prod}} \qquad \text{(R1)}$$

Ehn et al. (2014) also observed ELVOC formed from OH-oxidation of α-pinene with an estimated maximum molar yield of 1 %. In this work we evaluated the potential contribution of ELVOCs formed from OH-oxidation by using an ELVOC molar yield of 1 % for any monoterpene and sesquiterpene that reacts with OH (R2).

$$\mathrm{VOC}+\mathrm{OH}\rightarrow 0.01\mathrm{ELVOC}+0.99\mathrm{MCM_{ox.prod}} \qquad \text{(R2)}$$

In order to evaluate the potential influence of specific ELVOCs (e.g. dimers) which may be involved in the nano-CN formation (here denoted $ELVOC_{nucl}$) we also included the possibility to add separate reactions where the monoterpene and sesquiterpene ($VOC_{MT/SQT}$) that react with O_3 or OH form trace amounts of $ELVOC_{nucl}$ (R3-R4). In order to have little influence on the gas-phase chemistry or SOA formation, we used a very low $ELVOC_{nucl}$ yield (γ_{ELVOC}) of 0.001 mole % and scaled the nucleation rate coefficient accordingly.

$$\mathrm{VOC_{MT/SQT}}+\mathrm{O_3}\rightarrow \gamma_{\mathrm{ELVOC}}\mathrm{ELVOC_{nucl}}+\left(1\text{-}\gamma_{\mathrm{ELVOC}}\right)\mathrm{MCM_{ox.prod}} \qquad \text{(R3)}$$

$$\mathrm{VOC_{MT/SQT}}+\mathrm{OH}\rightarrow \gamma_{\mathrm{ELVOC}}\mathrm{ELVOC_{nucl}}+\left(1\text{-}\gamma_{\mathrm{ELVOC}}\right)\mathrm{MCM_{ox.prod}} \qquad \text{(R4)}$$

The reactions R3 and R4 were only used in those simulations where $ELVOC_{nucl}$ were involved in the nano-CN formation (Eq. 3-6).

3.2 Aerosol dynamics

The aerosol dynamics module in ADCHAM is based on the aerosol dynamics code from the

2-D Lagrangian model for Aerosol Dynamics, gas-phase CHEMistry and radiative transfer (ADCHEM) (Roldin et al., 2011). It includes subroutines for nano-CN formation, condensation/evaporation, Brownian coagulation and particle deposition onto the chamber walls. ADCHAM simulates the condensation, dissolution and evaporation of sulphuric acid, ammonia, nitric acid, hydrochloric acid and an unlimited number of organic compounds using the analytic prediction of condensation scheme (Jacobson, 1997) and prediction of non-equilibrium growth scheme (Jacobson, 2005). The aerosol particle water content is calculated with a thermodynamics model (Sect. 3.4).

ADCHAM considers the deposition of particles onto the chamber walls and keep track of the amount of deposited material on the walls. In Roldin et al. (2014) we kept track of each compound in each particle size bin that deposited on the chamber walls, and explicitly simulated the mass-transfer-limited gas-particle partitioning between the gas-phase and the wall deposited particles, assuming that the particles deposited on the walls remain as spherical particles on the walls. In this work, we instead assumed that the deposited SOA particles lose their individual particle identity and merge into the VOC wall matrix together with the gas-phase VOCs that deposit directly to the chamber walls (Sect. 3.5).

During the JPAC experiments the particles were formed by nucleation and, as a result of coagulation, wall losses and dilution, their average lifetime in the chamber was relatively short (less than 45 minutes, wherein dilution generally is the dominant loss process). Therefore, we expect that most of the formed particles were under charged with respect to the Bolzmann charge equilibrium distribution so that we do not have to consider the enhanced deposition rates induced by the particle charge (McMurry and Rader, 1985, Pierce et al., 2008 and Roldin et al., 2014). Thus, we calculated the first order wall deposition loss rate assuming that all particles were non-changed using the indoor deposition loss rate model from Lai and Nazaroff (2000), which accounts for different deposition loss rates on upward-, downward- and vertically facing surfaces. In the model by Lai and Nazaroff (2000) the fundamental (but unknown) parameter for the particle loss rate is the friction velocity, u*. In this work we found that with a relatively small value of u* (0.02 m s^{-1}), the model best captures the observed particle number and particle volume concentration loss rates after the UV-light is turned off. With this value of u* the deposition loss rate is equal to the observed ELVOC molecule loss rate measured by Ehn et al. (2014) when the particle (molecule) diameter is equal to 0.8 nm. We therefore used the value of u* = 0.02 m s^{-1} for all the model results presented in this work.

We evaluated six nano-CN formation parameterizations (Eq. 1-6) and compared them against base-case simulations with a fixed nano-CN formation rate (*J*). The first mechanism (Eq. 1) is sulphuric acid activation nucleation (Kulmala et al., 2006), Eq. 2 is used for kinetic nucleation of two H_2SO_4 molecules (McMurry and Friedlander, 1979), Eq. 3 is used for kinetic nucleation of one H_2SO_4 molecule and one $ELVOC_{nucl,}$ Eq. 4 is the nano-CN parameterization proposed by Riccobono et al., (2014) based on experiments in the CLOUD chamber, Eq. 5 represents a mechanism were single ELVOCs serve as nano-CN (Ehn et al., 2014), and Eq. 6 represent a kinetic type of nucleation mechanism with $ELVOC_{nucl}$.

$$J = A_1[\mathrm{H_2SO_4}] \quad (1)$$

$$J = K_2[H_2SO_4]^2 \quad (2)$$

$$J = K_3[\mathrm{H_2SO_4}][\mathrm{ELVOC_{nucl}}] \quad (3)$$

$$J = K_4[\mathrm{H_2SO_4}]^2[\mathrm{ELVOC_{nucl}}] \quad (4)$$

$$J = A_5[\mathrm{ELVOC_{nucl}}] \quad (5)$$

$$J = K_6[\mathrm{ELVOC_{nucl}}]^2 \quad (6)$$

A (s^{-1}) and *K* (cm^3 s^{-1}) in Eq. 1-3 and Eq. 5-6 are formation rate coefficients for activation type and kinetic type of nucleation, respectively. For Eq. 4 the formation rate coefficient has the unit cm^6 s^{-1}.

The composition of the nucleation clusters was chosen in order to match the respective new particle formation mechanism. The dry nano-CN volume was composed of equal mole fractions of H_2SO_4 and $ELVOC_{nucl}$ when we used Eq. 1, 3 or 5, equal mole fractions of H_2SO_4 and NH_3 when we used Eq. 2, and pure $ELVOC_{nucl}$ when we used Eq. 6. The molar mass of $ELVOC_{nucl}$ was assumed to be 500 g/mol. The equilibrium water content of the nano-CN clusters was calculated with the thermodynamics model. The dry particle size of the nano-CN was assumed to be 1.5 nm.

Both $ELVOC_{nucl}$ and H_2SO_4 were assumed to be non-volatile. The $ELVOC_{nucl}$ first order wall loss rate was 0.011 s^{-1} according to Ehn et al. (2014). The $ELVOC_{nucl}$ condensation sink was modelled explicitly with the condensation algorithm in ADCHAM. Chemical degradation of $ELVOC_{nucl}$ was not considered.

The $ELVOC_{nucl}$ involved in the nano-CN formation were assumed to be formed instantaneously after the first oxidation stage of monoterpenes and sesquiterpenes (R3-R4). With Eq. 3 and 4 we investigated six different sources of $ELVOC_{nucl}$; (i) as a product formed exclusively from the ozonolysis of endocyclic monoterpenes (α-pinene and Δ^3-carene), (ii) from the ozonolysis of all monoterpene and sesquiterpene, (iii) from the ozonolysis of sesquiterpenes, (iv) from the OH and O_3 oxidation of monoterpenes and sesquiterpenes, (v) when monoterpenes and sesquiterpenes react with OH, or (vi) when sesquiterpenes react with O_3 or OH. When using Eq. 5 and 6 we only considered the $ELVOC_{nucl}$ that was formed from monoterpenes and sesquiterpenes reacting with OH.

3.3 Size distribution structures

ADCHAM can be operated with the full-moving, fixed-sections or moving-centre particle size distribution approach (Roldin et al., 2011). In this work, we have tested both the fixed-sections approach and full-moving method using different numbers of size bins. The largest advantage of the full-moving approach compared to the fixed-sections approach is that it does not introduce any numerical diffusion problems during condensation/evaporation because it allows the particles to growth/shrink to their exact size. The main disadvantage is that new particle size bins need to be introduced when new particles are formed. In this work, we handled this by only introducing new particles formed by nucleation once every minute and at the same time as we added a new size bin for the freshly nucleated particles, we also removed the size bin containing the largest (oldest) particles. In order to not remove any particles within a time frame substantially longer than their average residence time in the reaction chamber, which was about 45 minutes, we used 400 size bins for the full-moving approach. Thus, with this method we kept track of the nucleated particles for 400 minutes (almost 9 times the average residence time in the chamber), before they were removed from the modelled size distribution. This method was compared with the fixed-section approach using a different number of size bins in the diameter (D_p) range 1.5 to 800 nm. Based on these tests we could conclude that the fixed-section approach was not suitable due to numerical diffusion if the number of size bins were less than 400 (Fig. S2 in the supplementary material). With the full-moving approach the particle volume (PV) and particle number concentrations (PN) were well conserved although the last size bin was removed once every minute (Fig. S3). Therefore, we decided to use the full-moving method to represent the particle number size distribution in this work.

One difficulty with the full-moving method is that the particle number size distribution needs to be mapped back onto a fixed diameter grid in order to illustrate it as a dPN/dlogD_p distribution. This grid needs to be relatively coarse in order to not get a “jumpy” graphical representation of the particle number size distribution. For this we used 50 size bins between 1.5 and 360 nm in diameter.

3.4 Particle phase chemistry and phase-state

ADCHAM includes a detailed particle-phase chemistry module, which is used to calculate the particle equilibrium water content, the particle acidity, nitric acid and hydrochloric acid equilibrium vapour pressures for each particle size bin, and the non-ideal interactions between organic compounds, water and inorganic ions using the activity coefficient model AIOMFAC (Zuend et al., 2008, 2011). In this work, we did not model the specific interactions between the organic and inorganic compounds but assumed a complete phase-separation of the inorganic- and organic particle phase. We used AIOMFAC to calculate the equilibrium water content in both the inorganic and organic particle phase and the individual compound activity coefficients. The organic compound activity coefficients were used when deriving the organic compounds equilibrium vapour pressures above each particle size (Sect. 3.6).

The particle phase chemistry module also contains subroutines that can be used to calculate organic salt formation, oligomerization and heterogeneous oxidation (Roldin et al., 2014). Recently, Shiraiwa et al. (2013) illustrated what peroxyhemiacetal (PHA) formation between organic compounds containing aldehydes and hydroperoxide functional groups may proceed fast and contribute to a large and rapid increase of the formed SOA mass during photooxidation experiments. In this work we evaluate if this type of heterogeneous dimer formation mechanism may explain the observed nano-particle growth during the JPAC experiment. For this we tested to use a constant value of the PHA formation rate (k_{PHA}) of 12 M^{-1} s^{-1} adopted from Shiraiwa et al. (2013). We also tested to model k_{PHA} as a parameterization of the sulphate particle mole concentration ($x_{\mathrm{S(VI)}}$), assuming that the PHA formation is acid catalysed by the co-condensing H_2SO_4 (Eq. 7).

$$k_{PHA} = B \cdot x_{\mathrm{S(VI)}} \ \ (\mathrm{M^{-1}s^{-1}}) \tag{7}$$

Here B is a constant, the value of which we varied in the range 1-500 M^{-1} s^{-1}.

In this work we used the kinetic-multi-layer module in ADCHAM for investigating weather the phase-state of the SOA particles might have influenced the evolution of the particle number size distribution. In order to do this, we divided each particle into three layers (an approximately monolayer thick surface layer of 0.7 nm, and two bulk-layers). We considered the two extreme conditions where the SOA particles either were considered to be completely liquid-like (no concentration difference between the surface and bulk layers) or solid-like (no molecule transport between the surface and the particle bulk layers). Still, this had only a minor effect on the modelled particle growth (SOA formation), but by treating the SOA as solid-like improved the agreement between the modelled and measured SOA particle volume decay when the UV-light was turned off (Fig. S4). Thus, in the simulations used to produce the results presented in Sect. 4, we treated the SOA particles as solid-like with the assumption that the molecule transport between the particle bulk and the particle surface-layer is relatively slow compared to the time scale it takes for the condensation to form a new monolayer thick surface layer. However, it is important to mention that the effect of the SOA phase-state cannot be unambiguously distinguished from the effect of the SOA volatility when only looking at total particle volume loss rates.

3.5 Reversible VOC wall loss

The JPAC reaction chamber was mixed with a Teflon fan with mixing times <2 min. The first order VOC wall loss rate to the chamber walls (k_w) was therefore governed by the molecular diffusion across the boundary layer near the chamber walls and by the uptake rate at the wall surface. According to McMurry and Grosjean (1985), the first order VOC wall loss rate for FEP Teflon films ($k_{w,\mathrm{FEP}}$) can be modelled with Eq. 8, which has two key parameters; the VOC wall mass accommodation coefficient (α_w) and the coefficient of eddy diffusion (k_e). Unfortunately neither α_w or k_e can be derived easily. Based on the observed wall losses of particles, McMurry and Radar (1985) estimated k_e to be 0.12 s^{-1} in a 60 m^3 FEP Teflon film chamber. Zhang et al. (2015) estimated k_e to be 0.075 s^{-1} and 0.015 s^{-1} in two not actively mixed FEP Teflon film chambers with volumes of 24 and 28 m^3.

$$k_{w,\mathrm{FEP}} = \frac{A_w}{V_{chamber}}\left(\frac{\alpha_w \bar{\upsilon}/4}{1+(\frac{\pi}{2})\alpha_w\bar{\upsilon}/\left(4\sqrt{(k_e D)}\right)}\right) \tag{8}$$

Here A_w is the chamber wall surface area, $V_{chamber}$ is the chamber volume, $\bar{\upsilon}$ is the mean thermal speed of the gas molecules and D is the molecular diffusion coefficient.

In the JPAC reaction chamber Ehn et al. (2014) observed ELVOC first order wall loss rates in the range 0.013-0.011 s^{-1}. By inserting a value of 0.011 s^{-1} for $k_{w,\mathrm{FEP}}$ in Eq. 8 and assuming that the surface wall uptake rate is not limiting the ELVOC wall loss rate (α_w>10^{-3}) we get a k_e of 4.2 s^{-1} for the JPAC reaction chamber. This value is substantially larger than what was estimated by Zhang et al. (2014) and McMurry and Radar (1985) and is probably because the JPAC reaction chamber was actively mixed and has a smaller volume.

From measurements in a 4 m^3 FEP Teflon chamber Kokkola et al. (2014) observed that for nopinone which has a pure-liquid equilibrium saturation vapour pressure (p_0) of 53 Pa, the gas-wall equilibrium was reached within a few minutes and $k_{w,\mathrm{FEP}}$ was ≥0.03 s^{-1}. The observations by Ehn et al. (2014) and Kokkola et al. (2014) indicate that gas-wall equilibration can be reached rapidly for both volatile and low volatile VOCs and that their uptake on the chamber walls is primarily limited by the transport to the chamber walls and not by surface uptake (α_w). In contrast, Zhang et al. (2015) observed that the gas-phase concentration of 25 different oxidized VOCs ($p_0 = 6\text{x}10^{-6}$ - 20 Pa) slowly decreased over more than 18 hours without reaching gas-wall equilibrium. These experiments were performed in a 24 m^3 FEP Teflon film chamber that was not actively mixed. Based on Eq. 8, Zhang et al. (2015) concluded that the VOC wall loss rate onto the FEB Teflon film walls was primarily limited by the surface uptake (α_w) and not by the molecule diffusion to the chamber walls. Zhang et al., (2015) also derived a parameterization of α_w as a function of the compounds pure liquid equilibrium saturation vapour pressure.

Based on the rapid ELVOC wall losses observed by Ehn et al. (2014), we assumed that the VOC losses to the JPAC reaction chamber walls was primarily governed by the molecule diffusion to the chamber walls and used a constant k_w equal to 0.011 s^{-1} for all condensable organic compounds. However, we also performed test simulations using Eq. 8 and the α_w parameterization from Zhang et al. (2015).

According to Matsunaga and Ziemann (2010) the loss rate of VOCs from FEP Teflon chamber walls back to the gas-phase (k_g (s^{-1})) can be represented by Eq. 9,

$$k_{g,i} = \frac{k_w}{(RT/p_{0,i}C_w/\gamma_{w,i})} \qquad (9)$$

where C_w in Eq. 9 is an effective mole concentration of organic compounds on the chamber walls (mol m^{-3}) and $\gamma_{w,i}$ is the activity coefficient of compounds *i* in the organic film on the chamber walls, *T* is the temperature in Kelvin and *R* is the universal gas constant (J mol^{-1} K^{-1}). According to Matsunaga and Ziemann (2010), the FEP Teflon film walls serve as a large organic pool where VOCs can absorb ($C_w/\gamma_{w,i}$ on the order of 100 μmol m^{-3}). Thus, the actual VOCs deposited on the chamber walls is considered to have a negligible contribution to C_w.

We assumed that there was practically no C_w on the glass for the VOCs to dissolve into. We also estimated that even if the deposited VOCs and SOA were distributed as a molecule monolayer on the walls, the VOCs formed during the simulated experiment could only cover maximum 2 % of the total glass wall surface area in the end of the measurement campaign (see supplementary material). Thus, we assumed that the VOC losses onto the glass wall surfaces could be treated as a condensation process but without dissolution (absorption) into an organic matrix on the walls. With this approach the uptake of compound *i* is governed by the difference between the concentration in the gas-phase ($c_{g,i}$) and the pure liquid equilibrium saturation concentration ($c_{0,i}$) (Eq. 10-11). Thus, as long as $c_{g,i} < c_{0,i}$ the VOC will not condense on the glass walls. For many of the semi-volatility organic compounds (SVOCs), the limit at which $c_{g,i}$ become larger than $c_{o,i}$ was never reached during the experiments.

$$\frac{dc_{g,i}}{dt} = -k_w\left(c_{g,i} - c_{0,i}\right) \qquad \left(if\ c_{g,i} > c_{0,i}\ or\ if\ c_{w,i} > 0\right) \tag{10}$$

$$\frac{dc_{w,i}}{dt} = k_w\left(c_{g,i} - c_{0,i}\right) \tag{11}$$

For the PTFE Teflon floor, we used the theory developed for FEP Teflon films (Eq. 9, 12 and 13), assuming that the PTFE Teflon surface (in total 15 % of the total chamber surface area) has a $C_w/\gamma_{w,i}$= 100 µmol m^{-3} and $k_{w,i}$ = 1.7x10^{-3} s^{-1} (0.15 x 0.011 s^{-1}) .

$$\frac{dc_{g,i}}{dt} = -k_{w,i}c_{g,i} + k_{g,i}c_{w,i} \tag{12}$$

$$\frac{dc_{w,i}}{dt} = k_{w,i}c_{g,i} - k_{g,i}c_{w,i} \tag{13}$$

In order to mimic the actual experiment procedure (Sect. 2), we started the model simulations by running ADCHAM 14 days prior to the actual experiments started, with conditions similar to Day-1 of the experiment campaign (see Sect. 3.7). This allowed the VOC concentration to build up on the chamber walls.

3.6 Condensable organic compound properties

In this work, we used three different approaches to simulate the SOA formation. In the first approach, the SOA formation was modelled by considering the gas-particle partitioning of all non-radical organic compounds from the MCMv3.2 gas-phase chemistry code with estimated pure-liquid equilibrium saturation vapour pressure lower than 10^{-2} Pa at T=289 K (in total 488 compounds). The pure-liquid equilibrium saturation vapour pressures of these compounds were estimated using the boiling point and vapour pressure extrapolation method from

Nannoolal et al. (2004; 2008), hereafter referred to as the Nannooal method. MCMv3.2 only includes one sesquiterpene: β-caryophyllene. As an attempt to take into account the SOA formation from the other sesquiterpens, exclusively in these simulations all the sesquiterpenes were assumed to be emitted as β-caryophyllene. In this work we will refer to this SOA formation representation as the MCM compound SOA formation mechanism.

The second method is the 2D-VBS approach (Jimenez et al., 2009; Donahue et al., 2011), which is a simplified approach to describe how the volatility distribution of the VOCs (as a function of their Oxygen to Carbon ratio (O:C)) evolves beyond the first oxidation step. A detailed description of the structure and general assumptions of the 2D-VBS method used in ADCHAM is given by Hermansson et al. (2014). Here we primarily describe how the 2D-VBS was modified in order to better capture the SOA formation observed during the JPAC experiments.

The 2D-VBS compounds were distributed across 11 volatility classes separated by powers of 10 in equilibrium saturation concentration (at the standard temperature 298 K (C^*_{298})), ranging from 10^{-7} to 10^{3} μg m^{-3}, and 12 discrete O:C from 0.1 to 1.2, in steps of 0.1. VOCs were introduced into the 2D-VBS by scaling the amount of reacted monoterpenes and sesquiterpenes with stoichiometric VBS yield parameterizations that represent the volatility distribution of the first generation oxidation products. We used the parameterizations from Table 3 in Henry et al. (2012). These parameterizations represent the first generation product volatility distributions formed during experiments where α-pinene was primarily oxidized by ozone (no OH-scavenger) or where α-pinene was oxidized by OH. We used the former parameterization to represent the first generation terpene (monoterpenes and sesquiterpenes) oxidation products during the UV-off periods and the later for the UV-light on periods. ELVOCs formed as first generation oxidation products from R1 were introduced into the 2D-VBS assuming that they had a O:C of 1 and C^*_{298} = 10^{-7} μg m^{-3}.

Hermansson et al. (2014) illustrated that the modelled SOA formation is sensitive to the volatility distribution and the assumed O:C of the first generation oxidation products. Moreover it depends on how the functionalization and fragmentation pattern is represented in the VBS and the reaction rate constant between the VOCs and OH. Equation 14 from Jimenez et al. (2009) was used to represent how the fraction of VOCs which fragmentize upon oxidation varies as a function of their O:C.

$$f_{frag} = (O:C)^{y} \tag{14}$$

Jimenez et al. (2009) assumed that y was equal to 1/6, that the 2D-VBS compounds were oxidized by OH with a reaction rate coefficient (k_{OH}) of 3x10^{-11} cm^3 s^{-1}, and that the compounds that fragmentize had equal probabilities to split at any of the carbon bonds. However, with this 2D-VBS setup, ADCHAM substantially overestimated the SOA formation (Fig. S5). This was mainly because a substantial fraction of the 2D-VBS fragmentation products were still less volatile than the reacting VOCs. If we instead assumed that all fragmenting compounds (on a mole fraction basis with equal proportions) end up into the VBS bins where C^*_{298} is at least 3 order of magnitude larger than the corresponding functionalization products, and decreased the probability of fragmentation somewhat (y=1/3), the modelled particle volume concentration agreed better with the measurements (Fig. S6). Thus, in the proceeding sections we used this fragmentation parameterization. We also performed sensitivity tests where we varied the value of k_{OH} in the range of 3×10^{-11} to 5×10^{-11} cm^3 s^{-1} and the O:C of the first generation oxidation products in the range of 0.3 to 0.5 (Fig. S6). Based on these tests, we decided to use the values of $k_{OH} = 5\times10^{-11}$ cm^3 s^{-1} and O:C = 0.4 for the first generation oxidation products in the simulations used to produce the results in Sect. 4.

The temperature dependence of the 2D-VBS equilibrium saturation concentrations were derived using the Clausius Clapeyron equation and assuming the following form for the heat of vaporization (ΔH; Epstein et al., 2010):

$$\Delta H = -11\times \log_{10} C^*_{300} + 129 \text{ kJ mol}^{-1} \qquad (15)$$

As the third approach we also tested to represent the SOA formation with a one-product model. With this method, only a single non-volatile condensable organic compound was used to represent the SOA formation. This compound was formed as a first generation oxidation product when any of the monoterpenes or sesquiterpenes reacted with O_3 or OH, assuming a molar yield of 25 %. The non-volatile condensable organic compound was assumed to have a molar mass of 325 g/mol. This simulation represents the limiting condition where the SOA formation is purely kinetically limited and not absorption partition limited, and it was also used for estimating how large the ELVOC yield would need to be if the particles were only growing by condensation of ELVOCs.

Both with the 2D-VBS and when the MCM compounds were used to represent the condensable organic compounds, the Raoult's law and the Kelvin equation were used to calculate the equilibrium vapour pressure ($p_{s,i,j}$) for each compound (i) above each particle size (j) (Eq. 16). With the 2D-VBS we assumed that the organic compounds were mixed ideally (unity activity coefficients (γ_i)). However, for the MCM compounds the molecule structure is known, so we calculated the organic compound activity coefficients in the SOA particle mixtures using AIOMFAC.

$$p_{s,i,j} = p_{0,i} x_{i,j} \gamma_{i,j} e^{\left(\frac{4M_i\sigma_i}{RT\rho_j D_{p,j}}\right)} \tag{16}$$

where $x_{i,j}$ is the mole fraction of compound i in the particle surface layer of particles in size bin j, M_i is the molar mass of compound i and ρ is the density of the absorbing phase. In this work the surface tension of the organic compounds (σ_i) was estimated to be equal to 0.05 N m^{-1} following Riipinen et al. (2010), even though we also tested the values of σ_i in the range 0.02 to 0.07 N m^{-1}. These sensitivity tests revealed that the modelled total particle number concentration decreased by approximately 10 % and the total particle volume concentration by approximately 30 % when the value of σ_i was increased from 0.02 to 0.07 N m^{-1} (Fig. S7).

3.7 Optimizing model input parameters

We used the ADCHAM model to simulate the nano-CN formation and growth during one measurement campaign conducted in the JPAC chamber. We chose to study the first four continuous days from the measurement campaign because these days had continuous measurements, coverage of the dataset was complete and UV light-on period was the same.

The measured temperature and relative humidity (RH) in the JPAC reaction chamber were used directly in the model as input. The values from the first four days are illustrated in Fig. 2. The purple bars in the figures illustrate the UV-on periods. The temperature stabilized around noon on Day-1, after that the temperature was ca. 16 °C for all days. The RH had minor fluctuations during the experiment. RH was kept to ca. 60% during UV-on periods for the first three days. On Day-4, the RH was around 55 %. In addition, 8 discharge lamps were used to simulate solar illumination on Day-1 in the plant chamber, but only 4 lamps were used on the remaining days to generate lower levels of VOC emissions from the trees.

Because the inflow of ambient air into the JPAC chamber was purified by an adsorption dryer, by default, the concentrations of NO, NO_2, SO_2, and CO in the inflow to the reaction chamber were assumed to be 0.03, 0.2, 0.015 and 15 ppbv, respectively. However, in order to evaluate how sensitive the model results was, in particular the modelled VOC composition and SOA formation, we also performed sensitivity tests where we varied the inflow concentrations of NO_x ($NO+NO_2$) within the range of 0.05-1 ppbv. Concentrations of VOCs in the inflow to the reaction chamber were based on the GC-MS and PTR-MS data. The initial concentrations of all VOCs were set to zero. The measured concentrations of isoprene, monoterpenes and sesquiterpenes (in the plant chamber) are illustrated in Fig. 3. The monoterpene and sesquiterpene concentrations displayed a certain diurnal distribution pattern, because the discharge lamps mimicked sunlight in the plant chamber. The total monoterpene concentration exceeded 4 ppbv on Day-1, and then decreased to a value below 2 ppbv on the following three days. In particular, the ocimene concentration was the highest on Day-1, which can be explained by the fact that ocimene emission is light dependent (Owen et al., 2002), and the simulated solar light intensity in the plant chamber from the discharge lamps was two times higher on Day-1 than during Day-2 – Day-4.

The isoprene concentration and the total sesquiterpene concentration were about an order of magnitude lower than the monoterpene concentration. Overall, the measured concentrations of terpenes were at similar levels as ambient air concentrations in boreal forest environments (Ruuskanen et al., 2009).

At the first attempt, we tried to simulate the O_3 and OH concentrations with the gas-phase chemistry mechanism. An OH sink has been determined in the empty chamber by direct OH-LIF measurements (Diss. Sebastian Broch) and has been always found to be within 2 - 4 s^{-1}. However, even with a OH sink of 4 s^{-1}, the model could not capture the observed gradual decrease in the O_3 concentration and increase in OH concentration during the UV-light on periods without either varying the UV-light intensity or the OH sink (Fig. 4).

Therefore, we decided to use the measured O_3 and H_2SO_4 concentrations, and estimated OH concentration, as input to the model for the simulations presented in Sect. 4. Figure 4a shows the measured O_3 concentration during the experiment, Fig. 4b shows the estimated OH concentration based on the observed loss rate of 2-butanol and Fig. 4c shows the measured H_2SO_4 concentration. When the UV-light was turned on the H_2SO_4 concentration increased very rapidly and had a distinct peak. This is most likely due to an initial peak in the OH concentration at the moment when the UV-lamps was turned on (indicated by the modelled OH concentration). In the model we represented this by setting the OH concentration to 5×10^7 molecules cm^{-3} for the 12 first minutes after the UV-light were turned on.

The photolysis rates were simulated by using the quantum yields and absorption cross sections reported at the MCMv3.2 web site. In order to evaluate the direct influence of the

254 nm wavelength UV-light on the VOC composition and SOA formation, we performed model simulations where the 254 nm UV-light was not considered when calculating the photolysis rates (*J_x)* for all compounds except O_3. These test showed that the modelled condensable organic compound composition, O_3 concentration (Fig. 4a), OH concentration (Fig. 4b) and the SOA formation were not significantly influenced directly by the 254 nm UV-light source, but strongly by the OH generated from the photolysis of O_3. Apart from O_3, the compound that was affected most by the 254 nm wavelength UV-light was H_2O_2 ($J_{\mathrm{H_2O_2}}$=1.26×10^{-4} s^{-1} with the UV-light source and $J_{\mathrm{H_2O_2}}$=3.72×10^{-8} s^{-1} without the UV-light source). Some of the VOCs containing carbonyl groups and nitrate groups were also influenced by the UV-light source. The most prominent change in the modelled photolysis rate in these two groups of VOCs occurred for glyoxal (a dialdehyde) ($J_{glyoxal}$=2.93×10^{-5} s^{-1} with the UV-light source and $J_{glyoxal}$=5.04×10^{-7} s^{-1} without the UV-light source) and methyl nitrate (CH_3NO_3) ($J_{\mathrm{CH_3NO_3}}$=5.89×10^{-5} s^{-1} with the UV-light source and $J_{\mathrm{CH_3NO_3}}$=2.51×10^{-8} s^{-1} without the UV-light source). These values can be compared with the rates at which glyoxal and CH_3NO_3 were oxidized by OH. At the OH concentration of 5×10^{7} molecules cm^{-3} (typical for the UV-light on periods) this rate was 5.00 ×10^{-4} s^{-1} for glyoxal and 1.08 ×10^{-6} s^{-1} for CH_3NO_3. Thus, for glyoxal the reaction with OH was still ~10 times faster than the photolysis reaction rate, while for CH_3NO_3 the photolysis reaction rate was ~2 times larger than the rate at which CH_3NO_3 was oxidized by OH. This indicates that short wavelength UV-light sources (as the one used in JPAC) may influence the VOC composition, especially if a large fraction of the VOCs contain nitrate functional groups (i.e. at high NO_x concentrations).

4 Results and discussion

Table 2 summarizes the different model tests that we performed in this work in order to constrain the VOC wall losses, the aerosol dynamics-, gas- and particle-phase chemistry mechanisms that can explain the nano-CN formation and growth observed during the JPAC experiments presented in Sect. 2.

4.1 Time series of BVOC concentrations

To investigate the potential contribution of BVOCs to the nano-CN formation and growth, it is essential to properly predict the time series of the BVOC concentrations. The modelled isoprene, monoterpene, and sesquiterpene concentrations in the JPAC reaction chamber are plotted in Fig. 5, together with isoprene and total monoterpene concentrations measured by the PTR-MS. The modelled isoprene concentrations were in a good agreement with the measurements during the UV-off period on Day-1. During the rest of the simulated four days period, the simulated isoprene concentration was generally lower than the observations. However, considering the low isoprene concentrations and the uncertainties in the PTR-MS measurements, we cannot draw any conclusions on weather the model actually underestimates the isoprene concentration or not.

The summation of the modelled monoterpene isomer concentrations reached the same level as the measurements. During the dark periods, the monoterpene concentrations in the reaction chamber decreased to about one third of their concentrations measured in the plant chamber. This was because of the in-flow dilution and chemical reactions with O_3. When the UV light was switched on, the monoterpene concentration decreased sharply due to OH oxidation, as seen both from the measurements and simulation results in Fig. 5. Before the UV-lights were turned on ~80 % of the reacting monoterpenes were oxidized by O_3. At the UV onset this number dropped to ~10 % and in the end of the UV-on periods only ~2 % of the oxidized monoterpenes were oxidized by O_3. The modelled monoterpene concentration shows a somewhat more pronounced decrease during the UV-on periods than measurement. However, as with the isoprene concentration, the relative uncertainties in the PTR-MS measurements increases at lower concentrations and part of the PTR-MS unity resolution mass peak that was interpreted as monoterpenes may also have had small contributions from other VOCs.

The sesquiterpene concentrations were below the detection limit of the PTR-MS. Our modelled results show that the sesquiterpenes were strongly oxidized by O_3 once the tree emission from the plant chamber entered the reaction chamber (Fig. 3). Even when the UV-lights were turned on, 40 to 60 % of the sesquiterpenes were oxidized by O_3 and the rest by OH. The sesquiterpene concentrations decreased rapidly to below 0.01 ppbv. In addition, the sesquiterpene concentrations showed a minor decrease during UV-on periods because of the oxidation by OH.

Although a large fraction of the monoterpenes and sesquiterpenes were oxidized in the JPAC reaction chamber, nano-CN formation was only observed during the UV-on periods. This indicates that the terpene oxidation products formed during the dark periods (primarily by ozonolysis) either did not have the right properties or were not abundant enough to form, activate and/or grow nano-CN to sizes above the detection limit of the PSM-CPC setup.

4.2 Simulations of the observed new particle formation events

The first simulations were designed in order to constrain the VOC wall losses and the mechanisms responsible for the observed particle growth (Sect. 4.2.1 and Sect. 4.2.2). For these simulations, we used a fixed nano-CN formation ratio of 20 cm^{-3} s^{-1} during the UV-light on periods, except for the first 12 minutes with UV-light on for which we used J = 80 cm^{-3} s^{-1}, for Day-1 and Day-2 and 60 cm^{-3} s^{-1} for Day-3. During the UV-light off periods the nano-CN formation rate was zero. The nano-CN had a dry diameter of 1.5 nm and was assumed to be composed of equal number of sulphuric acid and ELVOC$_{\mathrm{nucl}}$ molecules. After, constraining the VOC wall losses and the potential particle growth mechanisms, we investigated several nano-CN formation mechanisms (Eq. 1-6) and compared the results against the results from the simulations with a fixed nano-CN formation rate (Sect. 4.2.3).

4.2.1 Modelling the reversible VOC wall deposition

Figure 6a shows the modelled total particle volume concentration (PV) starting 14 days prior to the start of the intensive experimental campaign. The measured PV during the campaign is also displayed. The model results are from simulations with the 2D-VBS. As was described in Sect. 3.5, for the base case model simulation we used a constant species independent k_w equal to 0.011 s^{-1} and Eq. 9-13 to differentiate between the reversible VOC wall losses to the glass and PTFE Teflon surfaces. The results in Fig. 6b and c which are from the base case simulation illustrate how the 2D-VBS compounds with different $C*$ were distributed onto the glass walls and the PTFE Teflon floor, respectively. The LVOCs (C^*_{298}<10^{-1} µg m^{-3}) deposited both onto the glass and Teflon surfaces, while the SVOCs (C^*_{298}>10^{-1} µg m^{-3}) were exclusively found at the Teflon floor where they could absorb into C_w. Some of the compounds with C^*_{298} = 10^{-1} µg m^{-3} deposited onto the glass walls shortly after the UV-light was turned on, but when the condensation sink was increased and when the terpene inflow concentration into the reaction chamber was lowered (Day-2 – Day-4), the gas-phase concentrations of these compounds never reached above their pure liquid equilibrium saturation concentrations, so they evaporated from the glass walls. Figure S8 in the supplementary material shows similar results but for a simulation with the MCM compound SOA formation mechanism and additional ELVOCs (R1) contribution to the particle growth.

Figure 6a also shows the modelled particle volume concentration when assuming that the

JPAC reaction chamber walls behave similarly to FEP Teflon walls (Eq. 8-9), in which case k_e was 4.2 s^{-1} and the a_w parameterization was taken from Zhang et al. (2015). With this parameterization, $k_{w,i}$ varies from 2×10^{-5} to 7×10^{-4} s^{-1} for compounds with a molar mass of 300 g mol^{-1} and vapour pressures in the range 10^{-2} to 10^{-10} Pa. Thus, in these model simulations the ELVOC wall losses were about 15 times lower than what was observed by Ehn et al. (2014). Because of the lower VOC wall losses, the model overestimated the SOA formation by a factor of 3-4 for Day-1, a factor of 1.5-2 for Day-2 and Day-3 and a factor of 2-3 for Day-4, there was no gradual increase in the SOA formation due to re-evaporation of SVOCs from the walls and the correlation between the model and measured PV is substantially worse. This illustrates that the wall losses in the JPAC chamber cannot be treated in the same way as in FEP smog chambers.

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Figure 7 shows the 2D-VBS VOC composition for: (a) the SOA particles, (b) the gas-phase, (c) the VOCs on the glass walls and (d) the VOCs on the PTEF Teflon floor, at Day-3 and after 5 hours with UV-lights on. The bar at O:C = 1, $\log_{10} C^*_{298}$=-7 corresponds to the ELVOCs which were assumed to be formed from ozonolysis and OH-oxidation of monoterpenes and sesquiterpenes (molar yield of 7 % and 1 %, respectively). After 16.5 days of continuous experiments, 1.4 mg m^{-3} of SVOCs had deposited onto the PTEF Teflon floor according to the model simulations. The largest fraction of the wall deposited VOCs were first generation terpene oxidation products (O:C=0.4). A fraction of these SVOCs can re-evaporated and reacted with OH in the gas-phase. This explains why the modelled SOA formation was gradually increasing during the first 15 days (Fig. 6a), although all the other model conditions were identical. A similar pattern in the modelled SOA particle volume was also observed when the SOA formation was simulated with the MCM compound SOA formation mechanism (Sect. 3.6) (Fig. S8a). The ELVOCs (that are formed as first generation oxidation products (R1)) and the LVOCs (formed from OH-oxidation of the first generation oxidation products), primarily deposited onto the glass walls. However, in total the VOCs deposited onto the glass walls only made up 5 % of the total amount of VOCs on the chamber walls.

4.2.2 Evaluation of potential particle growth mechanisms

When using the 2D-VBS, the modelled SOA composition was dominated by LVOCs and SVOCs formed from second- and multi-generation OH-oxidation products. This was the case even if we considered that ELVOCs were formed as first generation products after the O_3-

and OH-oxidation of all monoterpenes and sesquiterpenes (Fig. 7a). The reason for this is the large ELVOC wall losses in the JPAC chamber and the small condensation sink during the new-particle formation events. During the UV-light on periods, the gas-phase 2D-VBS VOC composition was dominated by oxidized SVOCs formed by fragmentation of the first generation oxidation products (Fig. 7b).

In Fig. 8 we compare the modelled (a) total particle number concentration and (b) total particle volume concentration with the observations from the PSM-CPC and the SMPS. The model results are from simulations with the 2D-VBS with or without ELVOC formation from the ozonolysis and OH-oxidation of all monoterpenes and sesquiterpenes (R1-R2) (molar yield of 7 % and 1 %, respectively), as well as from a simulation with the MCM compound SOA formation mechanism (Sect. 3.6) (including ELVOCs from the ozonolysis of α-pinene and Δ^3-carene (R1)). For this simulation we also considered rapid peroxyhemiacetal dimer formation in the particle phase using Eq. 7 and $B = 200\ \mathrm{M^{-1}\ s^{-1}}$. Without consideration of this type of a rapid acid catalysed dimer formation process, the particle growth was substantially underestimated and almost no SOA was formed when we simulated the SOA formation with the MCM compound SOA formation mechanism (Fig. S8a). Shown in Fig. 8 are also the results from a simulation with a one-product model approach (Sect. 3.6). In the model, this non-volatile compound was formed as a first generation oxidation product from O_3 and OH oxidation of all monoterpenes and sesquiterpenes with a molar yield of 25 % (mass yield of almost 60 %).

In all the model simulations, the model overestimated the SOA particle volume formation during Day-1 but tended to underestimate, or to give values similar to the measurements, during the period Day-2 - Day-4. The best agreement between the model and measured particle volume concentration was found with the 2D-VBS method ($R^2 = 0.699$ with ELVOCs formation and $R^2 = 0.697$ without ELVOC formation), even though the particle volume concentration could nearly equally well be represented with the MCM compound SOA formation mechanism and acid catalysed PHA dimer formation ($R^2 = 0.672$). In the simulation with the non-volatile one product model, the agreement between the modelled and measured particle volume concentration was worse ($R^2 = 0.634$), which indicates that the particle growth cannot be purely explained by condensation of ELVOCs. This model simulation also illustrates that even if the particle growth was only be kinetically limited (not absorption partitioning limited) and if the yield of the formation of condensable organic compounds was be the same in all the days, the decrease in SOA volume concentration from Day-1 to Day-2 should have been more pronounced because of the substantially smaller condensation sink during Day-2 - Day-4 compared to Day-1. In the other model simulations, this effect was partly but not fully compensated by the re-condensation of SVOCs from the walls, which contributed relatively more to the SOA formation during Day-2 - Day-4 compared to Day-1. The only remaining explanation we can find to why the model gave too much SOA particle volume during Day-1 compared to Day-2 - Day-4 is that the BVOC composition was substantially different during Day-1. It may be that ocimene which reacts rapidly with O_3 may not form SOA to the same extent as e.g. α-pinene.

Figure S9 in the supplementary material compares the modelled total particle volume concentrations from simulations with B = 10 or 200, or using a constant PHA dimer formation rate of 12 M^{-1} s^{-1} based on the work by Shiraiwa et al. (2013). With B = 10 the modelled PV are in very good agreement with the measured PV for Day-1 but for the following days the model substantially underestimated the PV. With a constant PHA dimer formation rate of 12 M^{-1} s^{-1} the model gives 3 times higher PV than the measurements for Day-1 but gives reasonable PV formation for Day-2 – Day-4.

Figure 9 shows the modelled SOA volatility distribution as a function of particle size. The results are from a simulation with the 2D-VBS, including ELVOC formation from all monoterpenes and sesquiterpens oxidized by O_3 and OH (R1-R2). As expected, the smallest particles contained the largest mole fraction of ELVOCs because of the Kelvin effect and because the uptake of the SVOCs is not kinetically limited but limited by absorption into the particle surface layer.

Figure 10 shows the modelled and measured particle number size distributions at ½, 1, 2 and 5 hours of UV-lights on, for each day of the experimental campaign. The model results are from a simulation with the 2D-VBS including ELVOC formation from all the monoterpenes and sesquiterpenes oxidized by O_3 and OH (R1-R2). For Day-1, the model overestimated the particle growth rate, which can also be seen from the overestimated SOA formation (Fig. 8b). For the reminder of the experimental campaign, the modelled particle number size distributions were in good agreement with the SMPS measurements, except for the particles having a diameter <30 nm, for which the model gave substantially higher concentrations. A contributing explanation for this feature can be non-accounted diffusion losses of particles in the SMPS inlet.

In the supplementary we show how the volatility distribution of the MCM compounds and the SOA formation changed when the NO_x concentration in the inflow to the chamber was varied in the range 0.05 to 1 ppbv (Fig. S10). When the NO_x concentration was increased from 0.05 to 1 ppbv the particle SOA volume concentration was increased slightly (~10 %).

4.2.3 Evaluation of potential nano-CN formation mechanisms

In this section we evaluate the different nano-CN formation mechanisms described in Sect. 3.2. For these simulations we used the 2D-VBS to simulate the evolution of the condensable organic compounds. For all simulations in this section, ELVOCs were formed from all the monoterpenes and sesquiterpenes oxidized by O_3 and OH (R1-R2). Because the exact vapour pressures, formation mechanisms and concentrations of ELVOCs still are very uncertain, we cannot dismiss the possibility that the new particle formation (formation of particles with D_p>1.6 nm) was limited by the activation of nano-CN and not by the nano-CN formation rate itself. In this section we evaluate possible nano-CN formation mechanism with the assumption that it was the nano-CN formation that primarily limited the observed new particle formation during the experiments.

Table 3 gives the coefficient of determination (R^2) between the modelled total particle number concentration and measured total particle number concentration (D_p > 1.6 nm) from the PSM-CPC setup. The R^2 values are only given for those nano-CN mechanisms (parameterizations) that gave a R^2 >0.75. The R^2 values were calculated with the data from hour 11 to hour 96, for which the measured total particle number concentration from the PSM-CPC setup was complete. The largest R^2 values (>0.95) between the modelled and measured total particle number concentration were achieved with nano-CN formation mechanisms that involves both H_2SO_4 and ELVOCs. With Eq. 3, this were the case both if the $ELVOC_{nucl}$ molecules were formed from OH-oxidation or OH and O_3 oxidation of monoterpenes and sesquiterpenes. With Eq. 4, this was only the case if the $ELVOC_{nucl}$ molecules exclusively were formed from ozonolysis of sesquiterpenes. These are the only mechanisms for which the values of R^2 were higher than in simulations with a fixed nano-CN formation rate during the UV-light on periods. The relatively high R^2 value for the simulation with a fixed nano-CN formation rate during the UV-light on period is an indication that the variability in the dataset with respect to the compound(s) responsible for the nano-CN formation was too small to fully constrain the dominating nano-CN formation mechanism during the experiments. However, the dataset is still useful to reject the less likely nano-CN mechanisms and to narrow down the list of possible mechanisms. For this, the R^2 value from the simulation with the fixed nano-CN formation rate was used as a benchmark.

In Fig. 11 we compare the modelled total particle number concentration when calculating J using Eq. 1, 2, 3 and 4. With Eq. 3, the $ELVOC_{nucl}$ were either assumed to be formed from ozonolysis and OH-oxidation of monoterpenes and sesquiterpens, or only from OH-oxidation. With Eq. 4, the $ELVOC_{nucl}$ were formed from sesquiterpenes oxidized by OH. Displayed is also the result from the simulation with constant J = 20 cm^{-3} s^{-1} during the UV-light on periods and J = 0 cm^{-3} s^{-1} during the UV-light of periods. Based on Fig. 11 and the R^2 values, it is evident that the new particle formation could not be captured very well with H_2SO_4 activation (R1) or H_2SO_4 kinetic (R2) type of nano-CN formation. The simulation with R1 had a relatively high R^2 value but substantially overestimated the total particle number concentration during the UV-light off periods. The same is true for the kinetic type of nano-CN formation involving H_2SO_4 and $ELVOC_{nucl}$ formed from oxidation of the BVOCs with O_3 and OH. In this case, the model especially overestimated the total particle number concentration during the first day of the UV-light off period when the BVOC emissions were the highest. The best agreement between the modelled and measured total particle number concentration, both based on the R^2 value and Fig. 11, was achieved with the kinetic type of nucleation involving H_2SO_4 and $ELVOC_{nucl}$ (Eq. 3) formed from oxidation of monoterpenes and sesquiterpenes with OH exclusively.

5 Summary and conclusion

In this study, we used the ADCHAM model to simulate the nano-CN formation and growth during an experimental campaign in the Jülich Plant Atmosphere Chamber, which focused on new particle formation induced by photochemical reactions of VOCs emitted from real plants (Dal Maso et al., 2014). With the model we evaluated potential nano-CN formation and growth mechanisms and how the VOC chamber wall losses influenced the SOA formation and composition.

Our analysis on the JPAC experiments illustrates how complex it is to design smog chamber experiments that mimic the conditions during new particle formation events in the atmosphere, in addition to which it provides useful information about what need to be considered when designing such experiments. In this work we showed that the relative contribution of ELVOCs to the nano-CN formation and growth was effectively suppressed due to their rapid and irreversible wall losses and the relatively high OH concentrations during the UV-light on periods. With these respects, the conditions during the JPAC experiments were not directly comparable with typical conditions during new particle formation events in the atmosphere. In addition, the experimental conditions showed too small variability to fully constrain the dominating nano-CN formation and growth mechanisms. Despite these limitations, the JPAC experiments serve as a valuable dataset to narrow down the list of potentially important nano-CN formation and growth mechanism over the boreal forest region.

In the lack of ELVOCs in the gas-phase, one possible mechanism that can explain the initial particle growth is rapid heterogeneous dimer formation of SVOCs. In this work, we found out that acid catalysed peroxyhemiacetal formation between aldehydes and hydroperoxides may explain the observed particle growth. However, the particle growth could be modelled equally well if the SOA was formed by condensation of low-volatility second- and multi-generation OH gas-phase oxidation products simulated with a 2D-VBS approach. Thus, based on our model simulations, we cannot conclude weather the observed particle growth primarily was driven by low-volatility organic compounds formed in the gas-phase or by rapid dimer formation in the particle phase. However, without the later mechanism the model fails to reproduce the observed SOA formation when using the MCMv3.2 oxidation products as condensable organic compounds.

Our results suggest that H_2SO_4 is one of the key compounds involved in the new particle formation, but cannot solely explain the new particle formation process. During the simulated experiments, the best agreement between the modelled and measured total particle number concentration was achieved when using a nano-CN formation rate of the form $J = K[H_2SO_4][ELVOC_{nucl}]$, where $ELVOC_{nucl}$ was formed as an BVOC OH oxidation product.

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Table 1. The VOCs measured with GC-MS in the JPAC plant chamber. The measured concentrations were used as input for the ADCHAM model. The "other MTs" and "other SQTs" refer to other monoterpenes and sesquiterpenes than those specified in the table, respectively.

Isoprene				
α-pinene	β-pinene	myrcene	sabinene	camphene
ocimene	Δ^3-carene	α-terpinene	Δ-terpinene	α-phellandrene
β-phellandrene	terpinolene	tricyclene	other MTs	
farnesene	β-caryophyllene	α-longipinene	Δ-cardinene	other SQTs
2-butanol	hexanal	benzene	toluene	eucalyptol
nonanal	bornyl acetate	methyl salicylate		

Table 2. Summary of the model sensitivity tests that we performed in order to evaluate and constrain possible mechanisms for the formation and growth of nano-CN during the JPAC experiments.

Sensitivity test category	Varied parameters (method) 1	Varied parameters (method) 2
Condensable VOC properties method	MCMv3.2 + the Nannoolal vapour pressure method	2D-VBS
2D-VBS assumptions	Functionalization, fragmentation and OH reaction rates	O:C of the first generation products
Particle dry deposition loss rates	Varying u* in the range 0.1-0.01 m s^{-1}	
Gas-phase chemistry	Uncertainties related to the assumed NO_x inflow concentration	Influence of the UV-light on the VOC composition
Influence of ELVOCs on the particle growth	ELVOC formation from ozonolysis and OH-oxidation of terpenes	
Nano-CN formation and initial growth	Mechanisms (Eq. 1-6)	Surface tension 0.02-0.07 N m^{-1}
Oligomerization in the particle phase	Peroxyhemiacetal formation, possibly acid catalysed by co-condensing H_2SO_4 (Eq. 7)	
Non-ideal mixing in the particle phase	Activity coefficients from AIOMFAC	
Reversible VOC wall losses	With or without absorptive uptake on the glass walls	
SOA phase-state	Liquid-like SOA	Solid-like SOA

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Table 3. Coefficient of determination (R^2) between the modelled and measured (PSM-CPC) total particle number concentration during the experimental campaign.

Nano-CN formation mechanism	**$ELVOC_{nucl}$ source**	**R^2**
$J = 20\ cm^{-3}\ s^{-1}$ when UV is on		0.937
$J = K_2[H_2SO_4]^2$		0.787
$J = A_1[H_2SO_4]$		0.887
$J = K_3[H_2SO_4][ELVOC_{nucl}]$	MT and SQT ox. by O_3 and OH (R3, R4)	0.951
$J = K_3[H_2SO_4][ELVOC_{nucl}]$	MT and SQT ox. OH (R4)	0.955
$J = K_3[H_2SO_4][ELVOC_{nucl}]$	SQT ox. by O_3 and OH (R3, R4)	0.891
$J = K_4[H_2SO_4]^2[ELVOC_{nucl}]$	SQT ox. by O_3 and OH (R3, R4)	0.92
$J = K_4[H_2SO_4]^2[ELVOC_{nucl}]$	SQT ox. by O_3 (R3)	0.951
$J = A_5[ELVOC_{nucl}]$	MT and SQT ox. OH (R4)	0.815

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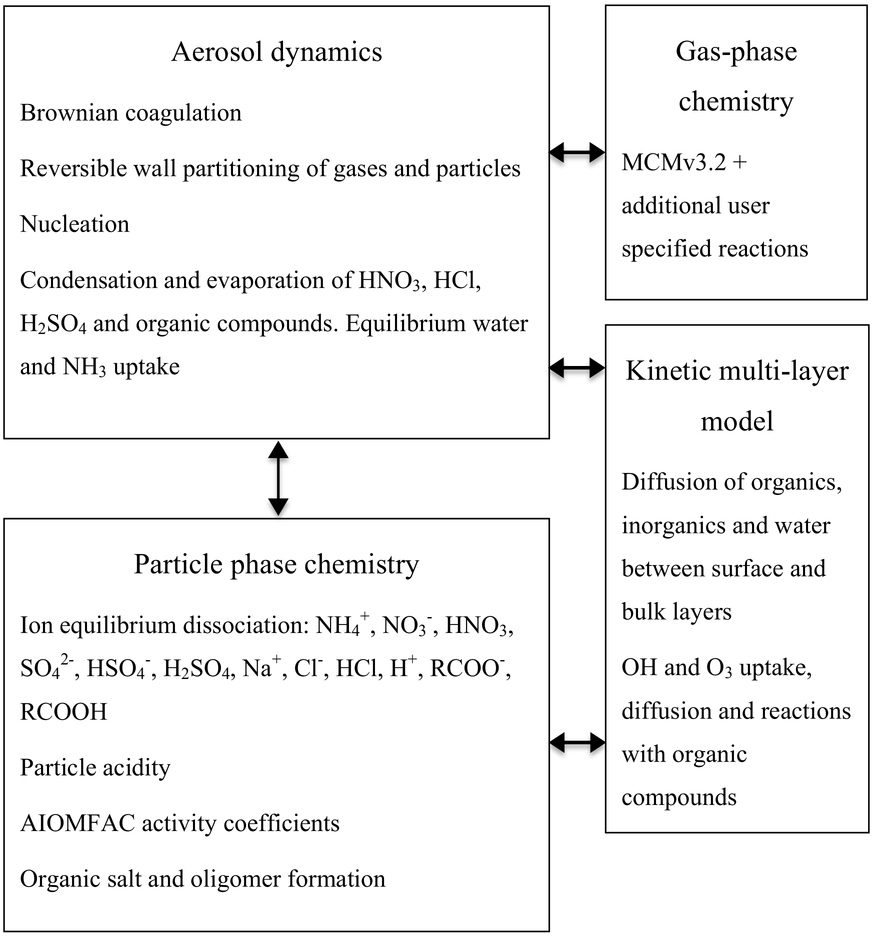

Figure 1. Schematic picture of the ADCHAM model structure.

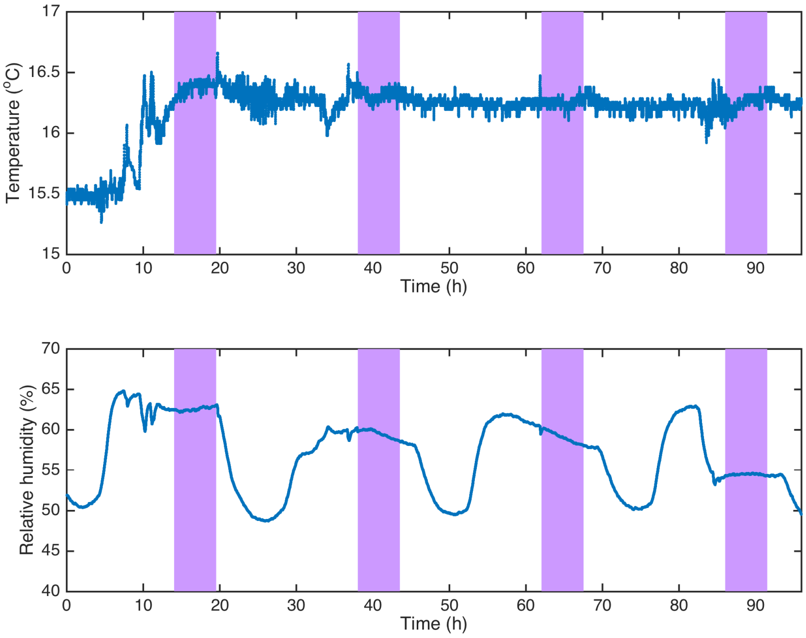

Figure 2. Measured temperature (a) and relative humidity (b) in the JPAC reaction chamber during the first four days of the measurement campaign. The purple bars indicate UV-on periods.

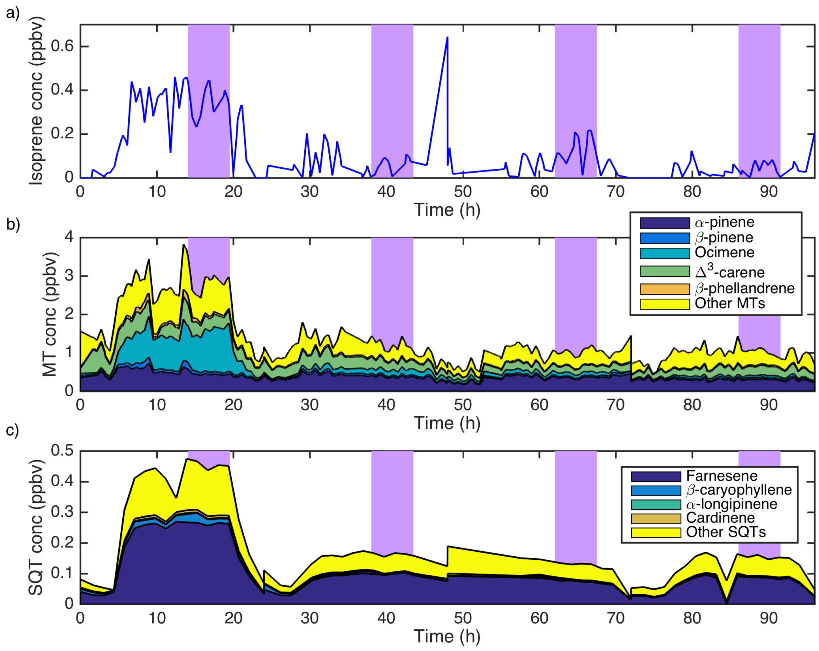

Figure 3. Isoprene (a), monoterpene (b) , and sesquiterpene (c) concentrations measured from the outlet air of the JPAC plant chamber. In Fig. b and c we left out those terpenes that had a negligible contribution to the total terpene concentration. Purple bars indicate UV-on periods during the measurements.

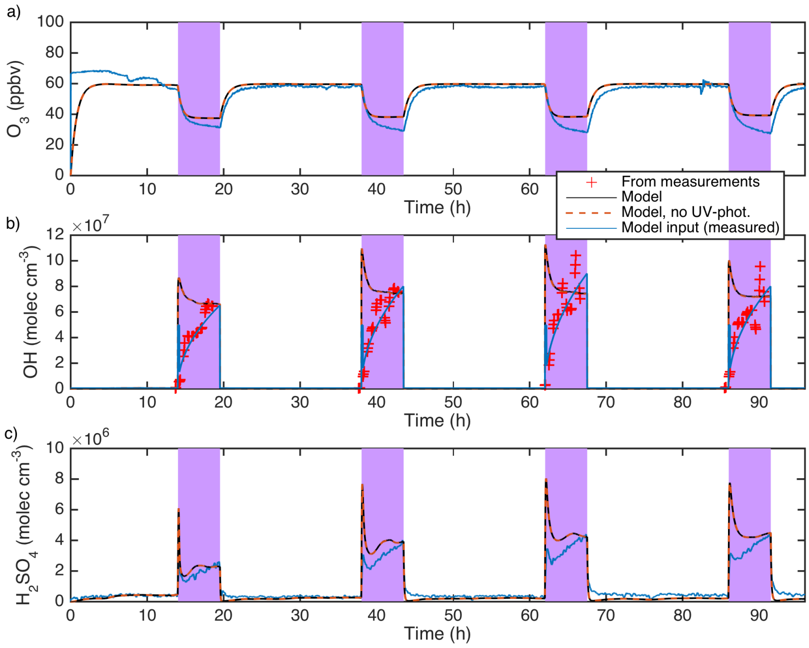

Figure 4. Measured and modelled (a) O_3 concentrations, (b) OH concentration, and (c) H_2SO_4 concentration. The OH concentration (red +) was not measured directly but was derived from the observed 2-butanol loss rate. The model results are from the simulations where the O_3, OH and H_2SO_4 concentrations were simulated by the model instead of given as the input concentration to the model. The results given by the orange dashed line corresponds to a simulation were the 254 nm UV-light was only used to calculate the photolysis rates of O_3 but not influencing photolysis rates of the other compounds. In the model simulations, we used the empirically determined JPAC reaction chamber-specific OH sink of 4 s^{-1} and an O_3 concentration of 170 ppbv in the inflow to the chamber. The blue lines give the model input concentrations that were used for the simulations presented in Sect. 4. For O_3 and H_2SO_4, the model input concentration was taken directly from the measurements.

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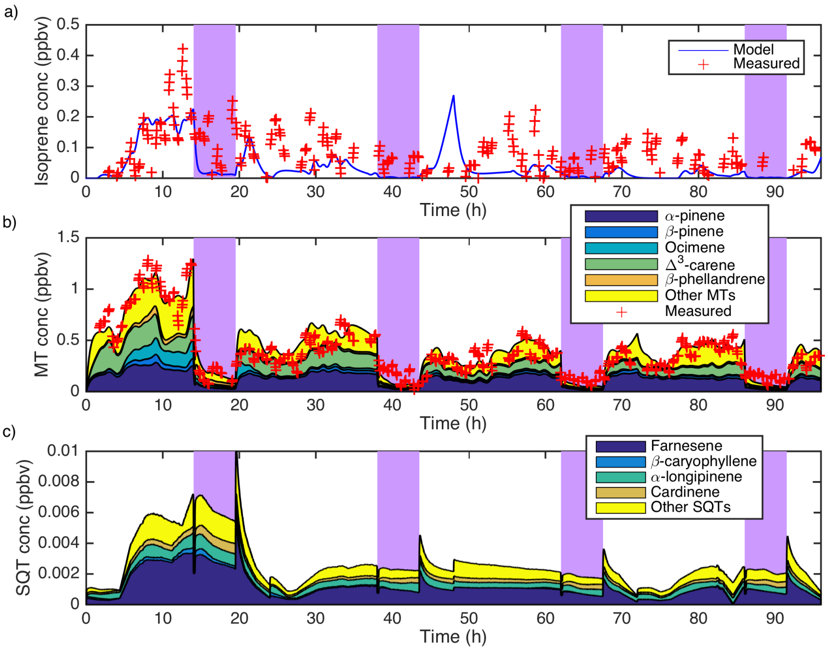

Figure 5. Modelled isoprene (a), monoterpene (b), and sesquiterpene (c) concentrations together with the measured isoprene and monoterpene concentrations in the JPAC reaction chamber. In Fig. b we left out those monoterpenes that had a negligible contribution to the total monoterpene concentration. The purple bars indicate UV-on periods during the measurements.

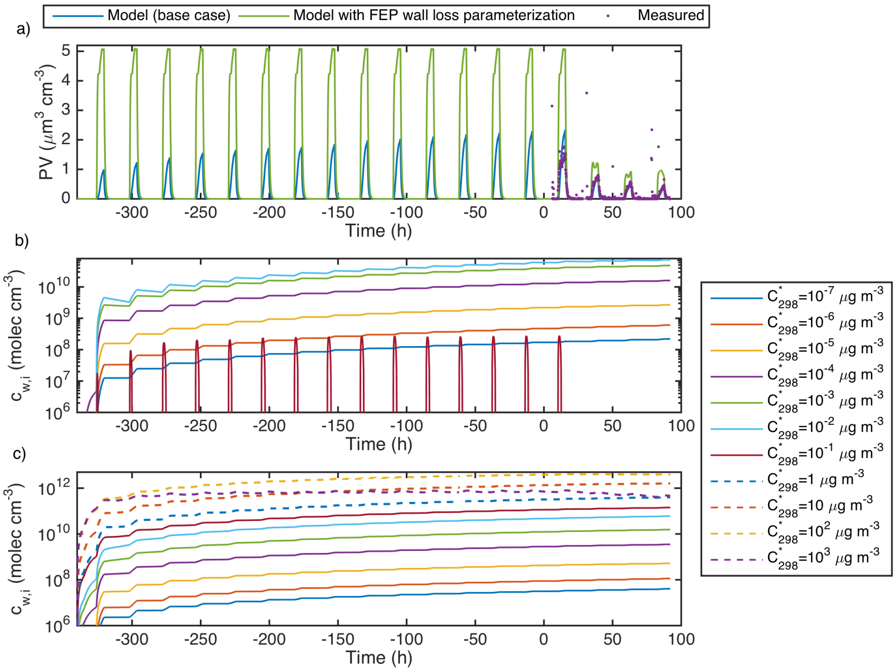

Figure 6. (a) Modelled and measured SOA volume concentration and 2D-VBS VOC wall uptake onto (b) the glass walls (Eq. 10-11) and (c) the PTFE Teflon walls when considering that the PTFE Teflon walls behave as FEB Teflon walls (Eq. 9, 12 and 13). The VOCs are summed over the all O:C but divided into the different C^*_{298} bins in the 2D-VBS. At time 0 h the intensive measurement campaign started. Fig 6a also shows the modelled SOA volume concentration when assuming that all chamber walls behave as FEP Teflon walls using Eq. 8-9, k_e = 4.2 s^{-1} and the VOC FEP Teflon wall mass accommodation coefficient parameterization from Zhang *et al*. (2015). ELVOC formation from R1-R2 was not considered for the simulations.

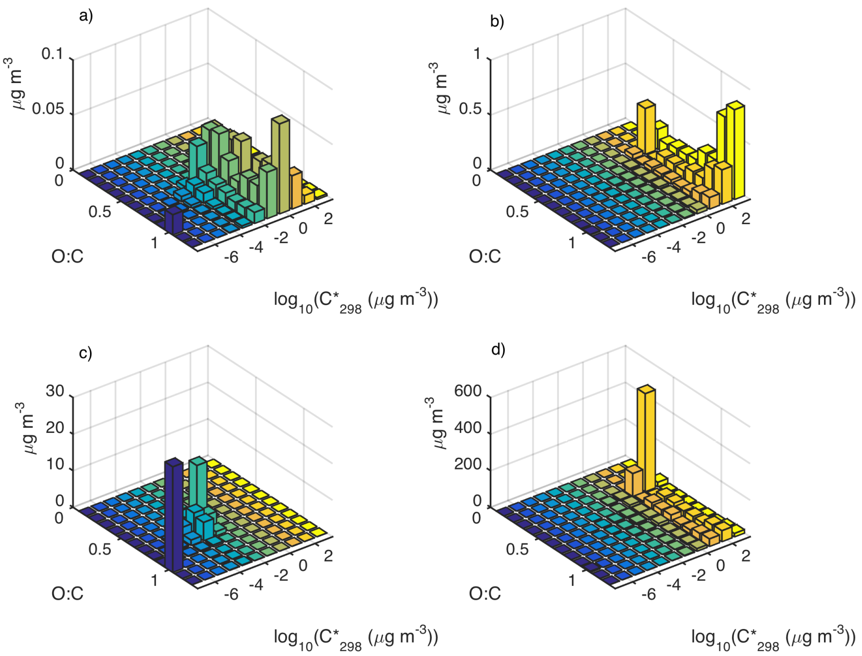

Figure 7. Modelled 2D-VBS distribution for (a) the SOA particles, (b) the gas-phase, (c) the VOCs on the glass-walls and (d) the VOCs on the Teflon floor. The model results are from Day-3 after 5 hours with UV-lights on. For this simulation R1 and R2 were used to simulate the ELVOC formation.

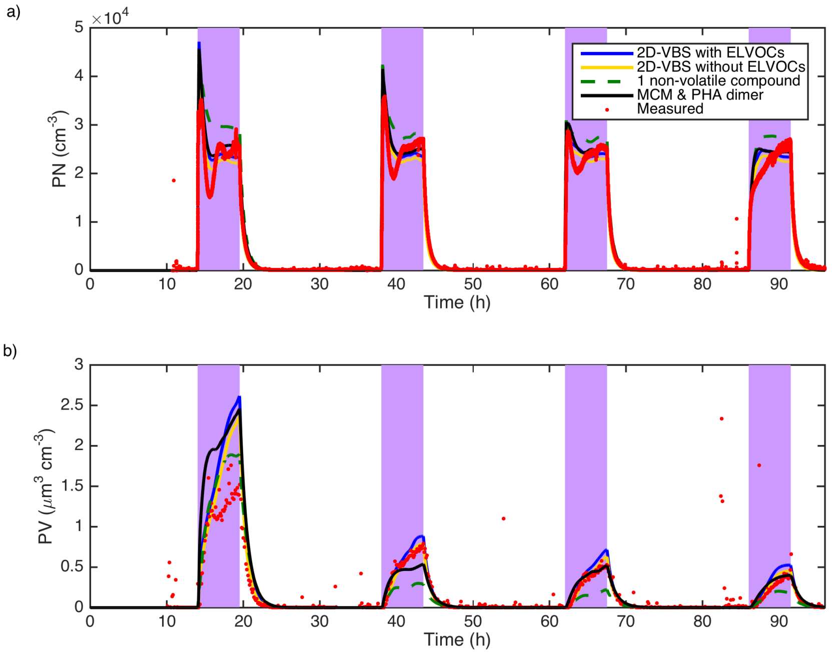

Figure 8. Modelled and measured (a) total particle number concentration and (b) total particle volume concentration. The model results are from simulations with the 2D-VBS and with or without ELVOC formation via ozonolysis and OH-oxidation of monoterpenes and sesquiterpenes (R1-R2), from a simulation when the MCM compounds were used as the condensable organic compounds and considering PHA dimer formation with Eq. 7, and from a simulation with only one condensable non-volatile compound. The mass yield of the non-volatile compound formed from O_3 and OH oxidation of all monoterpenes and sesquiterpens was 60 %.

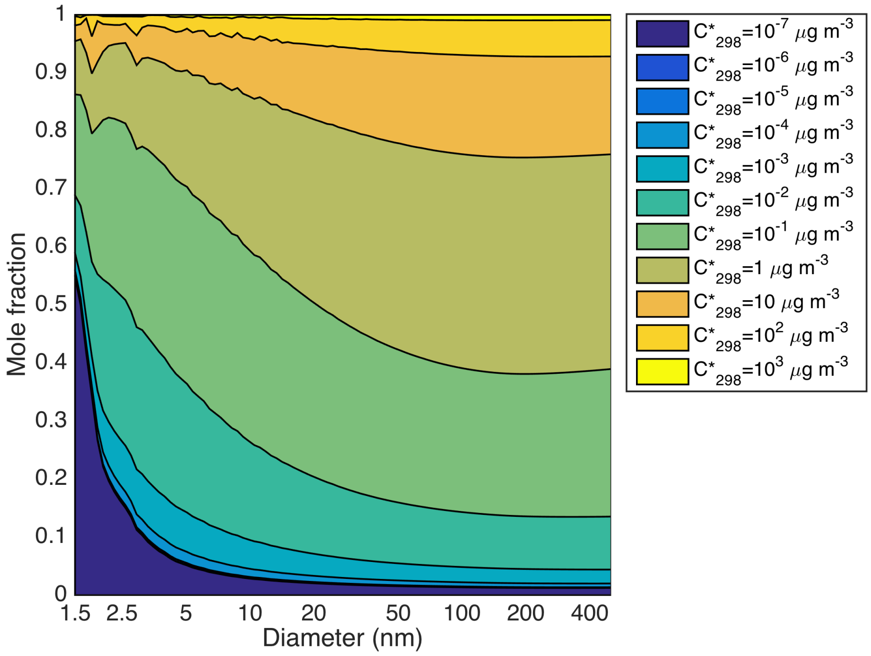

Figure 9. Modelled SOA volatility distribution as a function of the particle size from a simulation with the 2D-VBS and ELVOC formation from ozonolysis and OH-oxidation of monoterpenes and sesquiterpenes (R1-R2). The results are from Day-3 after 5 hours into the UV-light period.

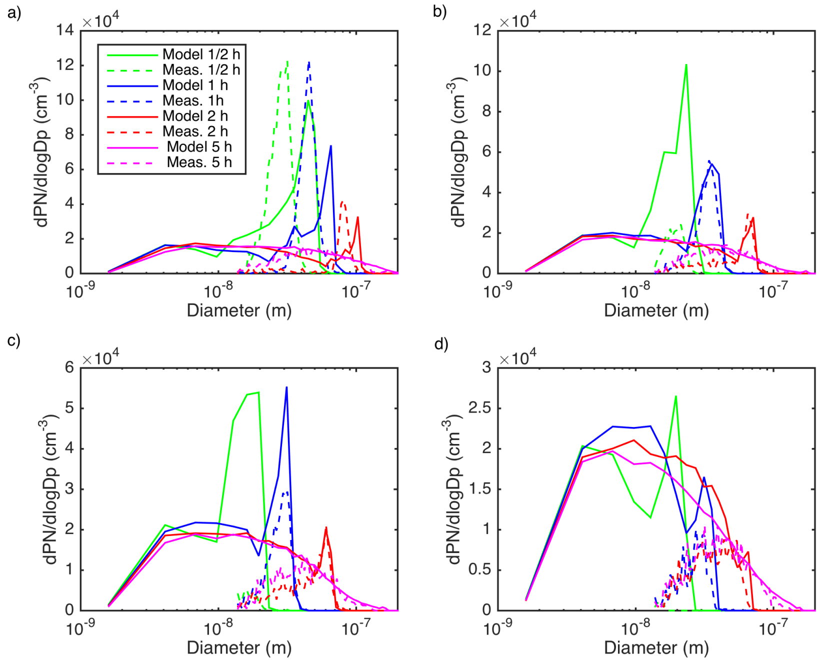

Figure 10. Modelled and measured particle number size distribution from (a) Day-1, (b) Day-2, (c) Day-3 and (d) Day-4 of the experimental campaign. The particle number size distributions are from ½, 1, 2 and 5 hours after the UV-light was turned on. The model results are from a simulation with the 2D-VBS (including ELVOC formation (R1-R2)) and the full-moving size distribution method.

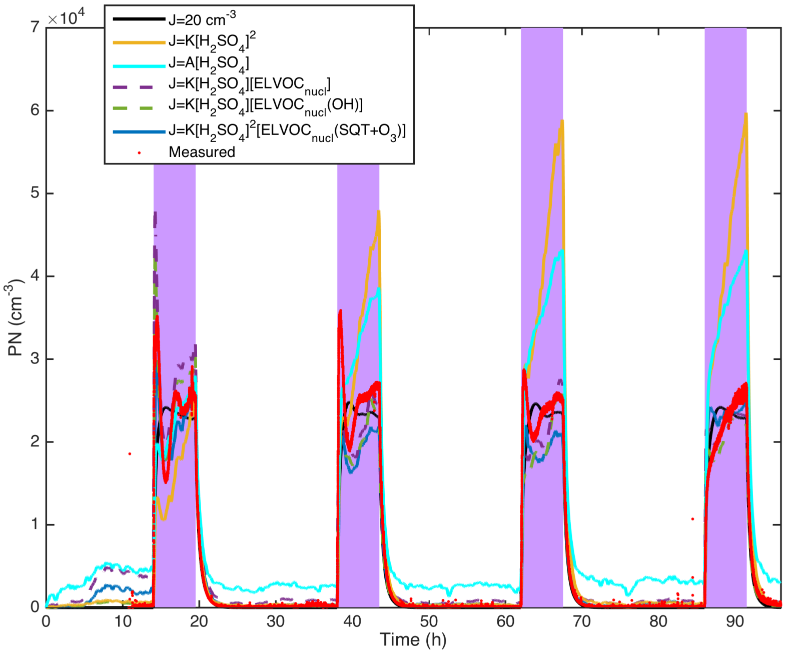

Figure 11. Measured (PSM-CPC) and modelled total particle number concentration with different nano-CN formation mechanism. The SOA formation was modelled with the 2D-VBS approach including ELVOC formation (R1-R4).