A note to editor and reviews regarding update in Supplementary Material:

Dear editor and reviewers,

Regarding the question by Reviewer #1 on the assumption of $Q_{s,ambient}$=$Q_{s,dry}$ in the $f(RH)$_water equation derivation, we responded that the $Q_{s,ambient}/Q_{s,dry}$ ratio was on average 1.40. We recalculated this ratio, considering variations in refractive index and using a broader range of particle sizes in the distribution, and found the ratio to be on average 1.06 for the SOAS study, and with a relative bias in $f(RH)$_water of 10%. This is discussed in the main body of the manuscript, and detailed in the Supplementary Material Section 2 accordingly.

~~Particle~~Fine particle water and pH in the southeastern United States

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~~Particle water and~~Fine particle pH in the southeastern United States

Abstract

Particle water and pH are predicted using ~~thermodynamic modeling (with ISORROPIA-II),~~ meteorological observations (RH, T), ~~and~~ gas/particle composition~~,~~ and thermodynamic modeling (ISORROPIA-II). A comprehensive uncertainty analysis is included, and the model is validated ~~with ammonia partitioning. The method is applied to predict~~. We investigate mass concentrations of particle water and related particle pH for ambient fine mode aerosols sampled in a relatively remote Alabama forest during the Southern Oxidant and Aerosol Study (SOAS) in summer~~,~~ and at various sites in the southeastern US~~,~~ during different seasons, as part of the Southeastern Center for Air Pollution and Epidemiology (SCAPE) study. Particle water and pH are closely linked; pH is a measure of the particle H^+ aqueous concentration~~,~~ and ~~so~~ depends on both the presence of ions and amount of particle liquid water. Levels of particle water, in-turn, are determined through water uptake by both the ionic species and organic compounds. ~~Particle ion balances, often used to infer pH, do not consider either the dissociation state of individual ions, nor particle liquid water levels and so do not necessarily correlate with particle pH.~~ Thermodynamic calculations based on measured ion concentrations can predict both pH and liquid water~~,~~ but ~~do not consider~~may be biased since contributions of organic species to liquid water ~~and so may also be biased.~~are not considered. In this study, contributions of both the inorganic and organic fractions to aerosol liquid water were considered and predictions were in good agreement with measured liquid water based on differences in ambient and dry light scattering coefficients (prediction vs. measurement: slope = 0.91, intercept = 0.~~45~~46 μg m^{-3}, ~~R~~R^2 = 0.~~87~~75). ISORROPIA-II predictions were ~~evaluated~~confirmed by ~~reproducing the observed gas-particle partitioning of NH_3.~~good agreement between predicted and measured ammonia concentrations (slope = 1.07, intercept = -0.12 μg m^{-3}, R^2 = 0.76). Based on this study, organic species on average contributed 35% to the total water, with a substantially higher contribution (~~63~~50%) at night. However, not~~The mean pH predicted in the Alabama forest (SOAS) was 0.94 ± 0.59 (median 0.93).~~ ~~Not~~ including contributions of organic water ~~has~~had a minor effect on pH (changes pH by 0.15 to 0.23 units), suggesting that predicted pH without consideration of organic water could be sufficient for the purposes of aqueous SOA chemistry. The mean pH predicted in the Alabama forest (SOAS) was 0.94 ± 0.59 (median 0.93). ~~).~~ pH diurnal trends followed liquid water and were driven mainly by variability in RH; ~~in~~during SOAS nighttime pH was near 1.5 ~~and during day~~, while daytime pH was near 0.5. pH ranged from 0.5 to 2 in summer and 1 to 3 in the winter at other sites. The systematically low pH levels ~~of predicted pH~~ in the southeast may have important ramifications, such as significantly influencing acid-catalyzed reactions, gas-aerosol partitioning, and mobilization of redox metals and minerals. Particle ion balances or molar ratios, often used to infer pH,

do not consider the dissociation state of individual ions or particle liquid water levels and so do not necessarily correlate with particle pH.

Keyword

Particle acidity, pH, particle water, LWC, f(RH), SOA, SOAS, SAS, SCAPE

1 Introduction

The concentration of the hydronium ion (H^+) in aqueous aerosols, or pH, is an important aerosol property that drives many processes related to particle composition and gas-aerosol partitioning ~~(Jang et al., 2002; Meskhidze et al., 2003; Gao et al., 2004; Iinuma et al., 2004; Tolocka et al., 2004; Edney et al., 2005; Czoschke and Jang, 2006; Kleindienst et al., 2006; Surratt et al., 2007; Eddingsaas et al., 2010; Surratt et al., 2010)~~(Jang et al., 2002; Meskhidze et al., 2003; Gao et al., 2004; Iinuma et al., 2004; Tolocka et al., 2004; Edney et al., 2005; Czoschke and Jang, 2006; Kleindienst et al., 2006; Surratt et al., 2007; Eddingsaas et al., 2010; Surratt et al., 2010). ~~Despite its importance, measuring~~Measurement of pH is ~~not conserved during aerosol sampling because dilution makes its measurement~~ highly challenging. ~~For this reason,~~ and so indirect proxies are often used to represent particle acidity. The most common is an ion balance~~,~~: the charge balance of measurable cations and anions (excluding the hydronium ion). Although correlated with an acidic (net negative balance) or alkaline (net positive balance) aerosol ~~(Surratt et al., 2007; Tanner et al., 2009; Pathak et al., 2011; Yin et al., 2014)~~(Surratt et al., 2007; Tanner et al., 2009; Pathak et al., 2011; Yin et al., 2014), an ion balance cannot be used as a ~~proxy for~~measure of the aerosol concentration of H^+ in air (i.e., moles H^+ per volume of air, denoted hereafter as H^+_{air}). This is due to two factors, first, an ion balance assumes all ions are completely dissociated, but multiple forms are possible, depending on pH (e.g., sulfate can be in the form~~,~~ of H_2SO_4, HSO_4^-, or SO_4^{2-}). ~~A comparison between an ion balance predicted H^+_{air} to that from the full analysis discussed below, for this data set, is shown in the Supplementary material.~~ Secondly, pH depends on the particle liquid water content (LWC~~) (~~), as pH is the concentration of H^+ in an aqueous solution~~), which~~. LWC can vary considerably over the course of a day and between seasons ~~and influence~~significantly influencing pH ~~(Seinfeld and Pandis, 2006)~~(Seinfeld and Pandis, 2006). Aerosol thermodynamic models, such as ISORROPIA-II ~~(Nenes et al., 1998; Fountoukis and Nenes, 2007)~~(Nenes et al., 1998; Fountoukis and Nenes, 2007) and AIM ~~(Clegg et al., 1998)~~(Clegg et al., 1998), are able to calculate LWC and particle pH, based on ~~measurements or model predictions~~concentrations of various aerosol species, temperature (T), and relative humidity (RH) and offer a more rigorous approach to obtain ~~the~~ aerosol pH ~~(Pye et al., 2013)~~(Pye et al., 2013). ISORROPIA-

II calculates the composition and phase state of an NH_4~~$^+$~~-~~SO_4~~$^{2-}$-NO_3~~$^-$~~-Cl~~$^-$~~-Na~~$^+$~~-~~Ca~~$^{2+}$-K~~$^+$~~-~~Mg~~$^{2+}$-water inorganic aerosol in thermodynamic equilibrium with water vapor and gas phase precursors. The model has been tested with ambient data to predict acidic or basic compounds, such as $NH_{3(g)}$, $NH_{4(p)}$~~$^+$~~, and $NO_{3(p)}$~~$^-$~~ (~~Meskhidze et al., 2003~~Meskhidze et al., 2003; ~~Nowak et al., 2006~~Nowak et al., 2006; ~~Fountoukis et al., 2009~~Fountoukis et al., 2009; Hennigan et al., 2014).

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LWC is a function of RH, particle concentration and composition, and ~~predicted to be~~is the most abundant particle-phase species in the atmosphere, at least 2-3 times the total aerosol dry mass on a global average ~~(Pilinis et al., 1995; Liao and Seinfeld, 2005)~~(Pilinis et al., 1995; Liao and Seinfeld, 2005). At 90% RH, the scattering cross-section of an ammonium sulfate particle can ~~be increased~~increase by a factor of five or more above that of the dry particle, due to large increases in size from ~~uptake of water (Malm, 2001).~~water uptake (Malm and Day, 2001). Because of this, LWC is the most important contributor to direct radiative cooling by aerosols ~~(Pilinis et al., 1995)~~(Pilinis et al., 1995), currently thought to be -0.45 Wm^{-2} (-0.95 Wm^{-2} to +0.05 Wm^{-2}) ~~(IPCC, 2013)~~(IPCC, 2013). LWC plays a large role in secondary aerosol formation for inorganic and possibly organic species by providing a large aqueous surface for increased gas uptake and a liquid phase where aqueous phase chemical reactions can result in products of lower vapor pressures than the absorbed gases~~.~~ (Ervens et al., 2011; Nguyen et al., 2013). In the eastern US, it has been suggested that the potential for organic gases to partition to LWC is greater than the potential to partition to particle-phase organic matter ~~(Carlton and Turpin, 2013)~~(Carlton and Turpin, 2013), and partitioning of water soluble organic carbon (WSOC) into the particle phase becomes stronger as RH (i.e., LWC) increases ~~(Hennigan et al., 2008)~~(Hennigan et al., 2008). Thus LWC enhances particle scattering effects directly by increasing particle cross sections (Nemesure et al., 1995) and indirectly by promoting secondary aerosol formation ~~(Nemesure et al., 1995)~~(Ervens et al., 2011; Nguyen et al., 2013).

The behavior of inorganic salts under variable RH is well established both experimentally and theoretically. It is known that dry inorganic salts (or mixtures thereof) exhibit a phase change, called deliquescence, when exposed to RH above a characteristic value. During deliquescence, the dry aerosol spontaneously transforms (at least partially) into an aqueous solution ~~(Tang, 1976; Wexler and Seinfeld, 1991; Tang and Munkelwitz, 1993)~~(Tang, 1976; Wexler and Seinfeld, 1991; Tang and Munkelwitz, 1993). In contrast, due to its chemical complexity that evolves with atmospheric aging, the relationship of organics to LWC is not well characterized~~, and requires a parameterized approach. Despite the abundance and importance of LWC, it is not routinely measured. Thus, actual ambient particle total mass concentration is not well characterized.~~ and requires a parameterized approach (Petters and Kreidenweis, 2007). Relationships between volatility, oxidation level and hygroscopicity are not always straightforward and still remain to be fully understood (Frosch et al., 2011; Villani et al., 2013; Cerully et al., 2014;

Hildebrandt Ruiz et al., 2014). Despite the abundance and importance of LWC, it is not routinely measured. Thus typically, particle total mass concentration (that includes liquid water) is often not characterized. In general, LWC is measured by perturbing the in-situ RH. The loss of particle volume when RH is lowered is assumed to be solely due to evaporated water. Approaches for LWC measurements are classified into single particle size probes and bulk size quantification ~~(Sorooshian et al., 2008)~~(Sorooshian et al., 2008). Single size particle probes provide more information, i.e., size resolved hygroscopic growth, and usually tend to be slow due to whole size range scanning. In contrast, bulk size measurements quantify the total water amount directly. The LWC measurement presented in this paper by nephelometers is a bulk measurement.

As part of the Southern Oxidant and Aerosol Study (SOAS), we made detailed measurements of particle organic and inorganic composition ~~(Xu et al., 2014), aerosol hygroscopicity~~ (Xu et al., 2015), aerosol hygroscopicity ~~(Cerully et al., 2014)~~(Cerully et al., 2014), and indirect measurements of particle LWC. These data are used to first determine the particle water mass concentrations, which are then utilized in a thermodynamic model for predicting pH. ~~This paper provides~~The fine particle LWC and pH data~~, which~~ from this analysis are used in our other studies of secondary aerosol formation as part of SOAS and discussed in companion papers to this work ~~(Cerully et al., 2014; Xu et al., 2015)~~(Cerully et al., 2014; Xu et al., 2015)~~. In this analysis we use bulk properties and do not consider variability in parameters with particle size. Particle phase separations are also not considered, although they have been measured in bulk extracts of aerosols from the southeast (You et al., 2012). With ambient RHs typically higher than 60%, formation of non-crystalline organic coatings over an aqueous core are possible, but glassy states are not expected. Good agreement between measured particle water and ammonia partitioning to predictions using the bulk properties (discussed below) suggests these assumptions are reasonable.~~.

2 ~~Date~~Data collection

2.1 Measurement sites

Aerosol measurements were conducted at the SEARCH ~~Centerville~~Centreville site (CTR; 32.90289 N, 87.24968 W, altitude: 126 m), located in Brent, Alabama, as part of ~~the~~ SOAS (Southern Oxidant and Aerosol Study)~~ study~~ (http://soas2013.rutgers.edu). SOAS ground measurements were made from June 1st to July 15th in the summer of 2013. CTR is a rural site within a large forested region dominated by biogenic ~~VOCs emission~~volatile organic compound (VOC) emissions, with minor local anthropogenic emissions and some plumes transported ~~plumes (sulfur, CO~~from other locations (coal-fired electrical generating units, urban emissions, biomass burning, mineral dust). It is representative of background conditions in the southeastern US and chosen to investigate biogenic secondary organic aerosol (SOA)

formation and its interaction with anthropogenic pollution transported from ~~long range transport~~other locations.

Additional measurements were also made at various sampling sites in and around the metropolitan Atlanta region from May 2012 to December 2012 as part of a large health study~~,~~; the Southeastern Center for Air Pollution and Epidemiology (SCAPE). A map of all ~~the~~ five sites is shown in Figure 1. The SCAPE measurement sites include:

1) A road-side (RS) site (33.775602 N, 84.390957 W), situated within 5m from the interstate highway (I75/85) in midtown Atlanta and chosen to capture fresh traffic emissions;
2) A near-road site (GIT site, 33.779125 N, 84.395797 W), located on the rooftop of the Ford Environmental Science and Technology (EST) building at Georgia Institute of Technology (GIT), Atlanta, roughly 30 to 40 m above ground level, 840 m from the RS site;
3) Jefferson Street (JST) (33.777501 N, 84.416667 W), a central SEARCH site representative of the Atlanta urban environment, located approximately 2000 m west of the GIT site;
4) Yorkville (YRK) (33.928528 N, 85.045483 W), the rural SEARCH pair of JST, situated in an agricultural region approximately 70 km west from the JST, GIT and RS sites.

More information on the SEARCH sites can be found ~~in~~elsewhere ~~(Hansen et al., 2003; Hansen et al., 2006)~~(Hansen et al., 2003; Hansen et al., 2006). We first focus on the SOAS campaign data, where ~~the widest~~wide range of instrumentation was deployed (http://soas2013.rutgers.edu~~),~~) to develop a comprehensive method of predicting LWC and pH, as well as assessing their uncertainties. The approach is then applied to the SCAPE site data to provide a broader spatial and temporal assessment of $PM_{2.5}$ ~~particle water and~~ pH in the southeastern US.

2.2 Instrumentation

2.2a PILS-IC~~:~~

$PM_{2.5}$ or PM_1 (particles with aerodynamic diameters < 2.5 or 1.0 μm at ambient conditions) water soluble ions were measured by a Particle-Into-Liquid-Sampler coupled ~~with~~to an Ion Chromatograph (PILS-IC; Metrohm 761 Compact IC). Similar setups are described in previous field studies ~~(Orsini et al., 2003; Liu et al., 2012)~~(Orsini et al., 2003; Liu et al., 2012). Metrosep A Supp-5, 150/4.0 anion column and C 4, 150/4.0 cation column (Metrohm USA, Riverside, FL) were used to separate the PILS liquid sample anions (sulfate, nitrate, chloride, oxalate, acetate, formate) and cations (ammonium, sodium, potassium, calcium, magnesium) at a 20 min duty cycle. The PILS sample ambient air flow rate was 16.8 ± 0.4 L min^{-1}. URG (Chapel Hill, NC) cyclones were used to provide PM cut sizes of $PM_{2.5}$ for the 1st half of field

study (June 1 to June 22) and PM_1 for the latter half (June 23 to July 15). Honeycomb acid (phosphoric acid)- and base (sodium carbonate)-coated denuders removed interfering gases before entering the PILS. The sample inlet was ~7 m above ground level and ~4 m long. The sampling line was insulated inside the trailer (typical indoor T was 25 ℃) and less than 1m in length to minimize possible changes in aerosol composition prior to measurement. Periodic 1~~-hour~~-hr blank measurements were made every day by placing a HEPA filter (Pall ~~corp~~Corp.) on the cyclone inlet. All data were blank corrected. The PILS IC was only deployed for the SOAS study.

2.2b AMS~~:~~

A High-Resolution Time-of-Flight Aerosol Mass Spectrometer (HR-ToF-AMS, Aerodyne Research Inc., hereafter referred to as "AMS") provided real time, quantitative measurements of the non-refractory components of submicron aerosols ~~(DeCarlo et al., 2006; Canagaratna et al., 2007)~~(DeCarlo et al., 2006; Canagaratna et al., 2007). In brief, particles were first dried (RH < 20%) and then immediately sampled through an aerodynamic lens into the high vacuum region of the mass spectrometer, then transmitted into a detection chamber where particles impact on a hot surface (600 ℃). Non-refractory species are flash vaporized and then ionized ~~with~~by 70 eV electron impact ionization. The generated ions are extracted into the time-of-flight mass spectrometer. Further details on the AMS setup and data processing can be found in Xu et al. (~~2014~~2015).

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2.2c CCNc~~:~~

The particle hygroscopic parameter, κ ~~(Petters and Kreidenweis, 2007)~~(Petters and Kreidenweis, 2007), used to infer the hygroscopic properties (liquid water associated with organics), was obtained from size-resolved CCN measurements ~~obtained with~~from a Droplet Measurement Technologies Continuous-Flow Streamwise Thermal Gradient Cloud Condensation Nuclei counter (CFSTGC, referred to hereafter as CCNc) ~~(Roberts and Nenes, 2005; Lance et al., 2006)~~(Roberts and Nenes, 2005; Lance et al., 2006). The CCNc exposes aerosols to a known supersaturation, then counts the activated particles that grow rapidly to droplet size. Theory can be used to parameterize the water phase properties (here, expressed by κ; (~~Petters and Kreidenweis (2007)~~Petters and Kreidenweis (2007)) of the organic aerosol, based on the size of particles that form CCN and their composition. A URG (Chapel Hill, NC) PM_1 cyclone was installed for both AMS and CCNc. The details of the CCNc setup and data analysis procedure can be found in ~~Cerully et al. (2014)~~Cerully et al. (2014).

2.2d Ambient vs Dry Nephelometers~~:~~

$PM_{2.5}$ (URG cyclones) aerosol light scattering coefficients (σ_{sp}) were measured online with two different nephelometers (Radiance Research M903) to infer LWC. Both were operated at nominally 3 L min^{-1} ~~controlled by critical orifices and vacuum pumps located some distance away from the instruments~~. Particle dry scattering was measured with a nephelometer located in the air-conditioned sampling trailer operated with a nafion dryer upstream~~, which~~ that maintained an RH of 31.5 ± 1.9 % (study mean ± SD, n = 12,464 based on 5-min averages). The other was situated in a small white 3-sided wooden shelter (one side covered by a loose tarp) located a distance away from all ~~building~~buildings to provide a scattering measurement at ambient T and RH. ~~Both~~Both $PM_{2.5}$ cut cyclones were located in ambient conditions, and both nephelometers were calibrated by CO_2 prior to the SOAS field campaign. Typical uncertainty is 3% for scattering coefficients ~~(Mitchell et al., 2009)~~(Mitchell et al., 2009). In addition, the nephelometer RH sensors were calibrated by placing the sensors in a closed container above aqueous saturated salt solutions that had reached equilibrium (measurements made in a thermally insulated container after a period of a few hours). Solution temperatures were monitored. Details on the calibration results are provided in the Supplementary ~~material.~~Material Section 1. Recorded RH was corrected by the ~~resulting~~ calibration results.

2.3 Determining LWC from nephelometers

Particle water was inferred from the ratio of ~~wet~~ambient and dry ~~PM_2~~$PM_{2.5}$ ~~(URG cyclones)~~ scattering coefficients (σ_{sp}) measured by the two nephelometers (defined here as aerosol hygroscopic growth factor, $f(RH) = $ ~~$\sigma_{sp(wet)}$~~$\sigma_{sp(ambient)}$ / $\sigma_{sp(dry)}$, where ~~$\sigma_{sp(wet)}$~~$\sigma_{sp(ambient)}$ and $\sigma_{sp(dry)}$ are particle scattering coefficients ~~in~~at ambient and dry RH conditions, respectively) following the method developed by other investigators ~~(Carrico et al., 1998; Kotchenruther and Hobbs, 1998; Carrico et al., 2000; Malm, 2001; Sheridan et al., 2002; Magi and Hobbs, 2003; Kim et al., 2006)~~(Carrico et al., 1998; Kotchenruther and Hobbs, 1998; Carrico et al., 2000; Malm and Day, 2001; Sheridan et al., 2002; Magi and Hobbs, 2003; Kim et al., 2006). A difference between ambient and dry scattering coefficients is assumed to be caused solely by loss of water. Detailed derivations are provided in the Supplementary ~~material~~Material. $f(RH)$ is related to the ~~measured~~ particle scattering efficiencies (~~Q~~Q_s) and average particle diameter ($\overline{D_p}$) by;

$$\overline{D_{p,wet}}\overline{D_{p,ambient}} = \overline{D_{p,dry}}\sqrt{f(RH)\,Q_{wet}/Q_{dry}}\sqrt{f(RH)\,\overline{Q_{s,dry}}/\overline{Q_{s,ambient}}} \quad (1)$$

~~Q_{wet}~~, $\overline{Q_{s,ambient}}$, ~~$\overline{D_{p,wet}}$~~ are$\overline{D_{p,ambient}}$ are the average scattering efficiency and average particle diameter under ambient conditions, while ~~Q_{dry}~~, $\overline{Q_{s,dry}}$, $\overline{D_{p,dry}}$ represent ~~the~~ dry ~~condition~~conditions. The method is based on ~~the fact that for~~ fine ~~particles~~particle light scattering ~~is most effective by~~being mostly due to

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particles in the accumulation mode and can be related to scattering efficiencies and the diameter of average surface, for both ~~wet~~ambient and dry particle size distributions. Assuming that $\overline{Q_{wet}}\overline{Q_{s,ambient}} = \overline{Q_{s,dry}}$ (see Supplementary Material Section 2 for justification and $\overline{Q_{dry}}$ ~~are of similar magnitude,~~uncertainty analysis), it follows then that;

$$\overline{D_{p,wet}}\overline{D_{p,ambient}} = \overline{D_{p,dry}}\sqrt{f(RH)} \quad (2)$$

Since the LWC is equal to the difference between ~~wet~~ambient and dry particle volume, we get;

$$LWC = [f(RH)^{1.5} - 1]m/\rho_p f(RH)\_water = [f(RH)^{1.5} - 1]m_p\rho_w/\rho_p \quad (3)$$

~~where m~~Where m_p and ρ_p are dry particle mass and density, respectively~~.~~; ρ_w is water density (constant 1 g cm^{-3} is applied). For SOAS, dry $PM_{2.5}$ mass concentrations were measured continuously by a TEOM (tapered element oscillating microbalance, 1400a, Thermo Fisher Scientific Inc., operated by Atmospheric Research & Analysis Inc~~.) were used.~~., referred to hereafter as ARA). Particle density, ρ_p, was computed from the particle composition, including AMS total organics, ammonium, and sulfate, which accounted for 90% of the measured $PM_{2.5}$ (TEOM) dry mass (SOAS study mean). A typical organic density 1.4 g cm^{-3} is assumed ~~(Turpin and Lim, 2001; King et al., 2007; Engelhart et al., 2008; Kuwata et al., 2012; Cerully et al., 2014)~~(Turpin and Lim, 2001; King et al., 2007; Engelhart et al., 2008; Kuwata et al., 2012; Cerully et al., 2014), and the density of ammonium sulfate is assumed to be 1.77 g cm^{-3} ~~(Sloane et al., 1991; Stein et al., 1994)~~(Sloane et al., 1991; Stein et al., 1994). ~~Particle density, ρ_p,~~ was calculated to be 1.49 ±0.04 g cm^{-3} (n = 4,393) using mass fractions (ε):

$$\rho_p = \frac{1}{\varepsilon(NH_4 + SO_4)/1.77 + \varepsilon\,(Organics)/1.4} \frac{1}{\varepsilon(NH_4^+ + SO_4^{2-})/1.77 + \varepsilon\,(Organics)/1.4}\ (g\ cm^{-3}) \quad (4)$$

The time-resolved composition data shows that dry particle density did not have a significant diurnal variability (±2.7%, SD/mean~~). (Note, SO_4 stands for sulfate in all its possible forms, from free to completely dissociated, the same applies to all other ions, e.g., NH_4, NO_3, etc.)~~, Supplementary Material Figure S2). In the following we refer to the particle water calculated by this method as $f(RH)$_water. The uncertainty of $f(RH)$_water is estimated to be ~~23%, mainly caused by Q_{wet}/Q_{dry} (~20%), m (10%), $\sigma_{sp(wet)}$~~15%, mainly caused by the calculation of $\overline{Q_{s,ambient}}/\overline{Q_{s,dry}}$ (LWC error of 10% from assuming $\overline{Q_{s,ambient}}/\overline{Q_{s,dry}}$ = 1, see Supplementary Material), m_p (10%), $\sigma_{sp(ambient)}$ / $\sigma_{sp(dry)}$ (4.2%) (uncertainty for a single σ_{sp} measurement is 3%, ~~Mitchell et al. (2009)~~Mitchell et al. (2009)~~), and ρ_p (2.7%).~~), and ρ_p (2.7%). Note that LWC error depends on RH, and for SOAS average composition aerosol could increase to 21% for RH > 90% (Supplementary Material Figure S6).

3 Modeling Methods: Predicting LWC ~~prediction~~and pH from aerosol composition

In most studies, such as SCAPE, particle water was not measured and must be determined based on aerosol composition. Both inorganic and organic components contribute to uptake of water vapor, establishing equilibrium for the ambient RH and T conditions. Thus, LWC is controlled by meteorological conditions and also by aerosol concentration and composition. Thermodynamic models, such as ISORROPIA-II, have been extensively used to predict LWC due to inorganic aerosol components ~~(Fountoukis and Nenes, 2007)~~(Fountoukis and Nenes, 2007). Contributions to LWC by organic components are typically based on an aerosol hygroscopicity parameter, κ, which is determined by CCN data. Here we refer to particle water associated with inorganics and organics as W_i and W_o, respectively. Total particle water ($W_i + W_o$) is taken as the sum of water associated with individual aerosol chemical components (sum of ions and lumped organics~~),~~) based on Zdanovskii-Stokes-Robinson (ZSR) relationship (Zdanovskii, 1936; Stokes and Robinson, 1966), with the assumption that the particles are internally mixed.

3.1 LWC from inorganic species

~~W_i was~~Particle water associated with inorganic species (W_i) were predicted by ISORROPIA-II ~~(Nenes et al., 1998; Fountoukis and Nenes, 2007)~~(Nenes et al., 1998; Fountoukis and Nenes, 2007). ISORROPIA-II calculates the composition and phase state of ~~an~~ a K^+-Ca^{2+}-Mg^{2+}-NH_4-~~SO_4-~~$^+$-Na^+-SO_4^{2-}-NO_3-~~-~~Cl-~~Na-Ca-K-Mg-~~-water inorganic aerosol in thermodynamic equilibrium with gas phase precursors. Chemical and meteorological data are necessary inputs. For our analysis at CTR, the inputs to ISORROPIA-II are the inorganic ions measured by the IC or AMS, RH measured by the outside ~~Nephelometer~~nephelometer, and temperature from the SEARCH site (ARA) meteorological data.

3.2 LWC from organic fraction

To determine the contributions to particle water by W_o, in SOAS the organic hygroscopicity parameter (κ_{org}) was ~~measured~~calculated based on the observed CCN activities of the organic fraction ~~(Cerully et al., 2014)~~(Cerully et al., 2014) ~~and in~~. In the following analysis diurnal three-hour running averages are used in the calculation. (Diurnal plot is included in the Supplementary ~~material~~Material as Figure ~~5)~~S7). W_o is calculated using the following equation ~~(Petters and Kreidenweis, 2007)~~(Petters and Kreidenweis, 2007).

$$~~W_o = \frac{m_s}{\rho_s}\frac{\kappa_{org}}{(1/RH - 1)}~~ W_o = \frac{m_{org}\rho_w}{\rho_{org}}\frac{\kappa_{org}}{\left(1/RH - 1\right)} \quad (5)$$

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~~where m_s is the organic mass concentration from AMS (Xu et al., 2014), and a typical organic density (ρ_s) of 1.4 g cm^{-3}~~ Where m_{org} is the organic mass concentration from AMS (Xu et al., 2015), ρ_w is water density, and a typical organic density (ρ_{org}) of 1.4 g cm^{-3} is used ~~(Turpin and Lim, 2001; King et al., 2007; Engelhart et al., 2008; Kuwata et al., 2012; Cerully et al., 2014)~~(Turpin and Lim, 2001; King et al., 2007; Engelhart et al., 2008; Kuwata et al., 2012; Cerully et al., 2014).

3.3 pH prediction

The thermodynamic model, ISORROPIA-II ~~(Fountoukis and Nenes, 2007)~~(Fountoukis and Nenes, 2007), ~~is used to predict particle pH, based on calculated~~ calculates the equilibrium particle hydronium ion concentration ~~in the aerosol.~~per volume air (H^+_{air}), which along with the LWC is then used to predict particle pH. To correct for the LWC associated with the organic aerosol (not considered in ISORROPIA-II), we recalculate pH by considering ~~the predicted particle hydronium ion concentration per volume air (H^+_{air})~~ H^+_{air} and total predicted water (W_i and W_o). The modeled concentrations are μg m^{-3} air for H^+_{air} and LWC. The pH is then,

$$pH = -\log_{10} H^+_{aq} = -\log_{10} \frac{1000 H^+_{air}}{W_i + W_o} \quad (6)$$

~~where~~Where H^+_{aq} (mol L^{-1}) is hydronium concentration in ~~a~~an aqueous solution.

ISORROPIA-II has been tested in previous field campaigns where a suite of both gas and particle components were measured ~~(Nowak et al., 2006; Fountoukis et al., 2009)~~(Nowak et al., 2006; Fountoukis et al., 2009). The model was able to predict the equilibrium partitioning of ammonia ~~(Nowak et al., 2006)~~(Nowak et al., 2006) in Atlanta and nitric acid ~~(Fountoukis et al., 2009)~~(Fountoukis et al., 2009) in Mexico City within measurement uncertainty. For instance, $NH_{3(g)}$, $NH_{4(p)}^+$, $HNO_{3(g)}$, and $NO_{3(p)}^-$ were within 10%, 20%, 80%, and 20% of measurements ~~(Fountoukis et al., 2009)~~(Fountoukis et al., 2009). In this study, ISORROPIA-II was run in the "Forward mode" for metastable aerosol. Forward mode calculates the equilibrium partitioning given the total concentration of various species (gas + particle) together with RH and T as input. Reverse mode involves predicting the thermodynamic composition based only on the aerosol composition. Here~~,~~ we use the Forward mode with just aerosol phase data input because it is less sensitive to measurement error than the Reverse mode ~~(Hennigan et al., 2014).~~(Hennigan et al., 2014). The W_i prediction ~~stays~~remains the same (Reverse vs Forward: slope = 0.993, intercept = -0.005, and R^2 = 0.99) no matter which approach is used. Gas phase input does have an important impact on the H^+_{air} calculation. ISORROPIA-II was tested with ammonia partitioning~~. Discussed~~, which is discussed in more detail below~~, here~~. Here it is noted that we found that further constraining ISORROPIA-II with measured $NH_{3(g)}$ ~~(You et al., 2014)~~(You et al., 2014) resulted in a pH increase of 0.8

at CTR and that the predicted $NH_{3(g)}$ matched the measured $NH_{3(g)}$ well (slope = 1.07, intercept = -0.12 μg m^{-3}, R^2 = 0.76). This also confirms that ISORROPIA-II predicts the pH in the ambient aerosol with ~~good~~reasonable accuracy, as inputting the total (gas + aerosol) ammonium results in predictions that agree with those observed. This is also in agreement with findings of ~~Hennigan et al. (2014)~~Hennigan et al. (2014) and ~~Fountoukis et al. (2009)~~Fountoukis et al. (2009), both of whom found that ISORROPIA-II reproduced the partitioning of ammonia and inorganic nitrate in Mexico City during the MILARGO campaign.

3.4 Assumptions

In the following analysis we use bulk properties and do not consider variability in parameters with particle size. Particulate organic and inorganic species are assumed to be internally mixed in the liquid phase due to the high RH (73.8 ± 16.1%) typical of this study and because a large fraction of the ambient aerosol organic component is from isoprene SOA (Xu et al., 2015), which are liquids at RH ≥ 60% (Song et al., 2015). Particle liquid phase separations are not considered, although they have been measured in bulk extracts of aerosols from the southeast (You et al., 2012). It is reported that liquid-liquid phase separation can occur when the O:C ratio of the organic material is ≤ 0.5. More experiments showed that it is possible to have phase separation for O:C ≤ 0.7, but not for O:C ≥ 0.8 (Bertram et al., 2011; Song et al., 2012; You et al., 2013). SOAS average O:C = 0.75 (± 0.12) is in the transition between these two regimes. According to Figure 2 in Bertram et al. (2011), at RH typically > 60% and organic:sulfate mass ratio >1, it is not possible to have phase separation, which is the case for our sampling sites. Based on our basic assumption of no liquid-liquid phase separation, pH is considered to be homogeneous in a single particle. However, separated phases would likely have different pHs if liquid-liquid phase separation occurs. In that case, pH should be calculated based on the amounts of water and H^+_{air} in each phase. Gas-particle phase partitioning will change according, due to these separated phases. There are models that are set up to calculate these thermodynamics (e.g., AIOMFAC), but none is yet able to address the compositional complexity of ambient SOA. (Zuend et al., 2010; Zuend and Seinfeld, 2012) Although it is often true that non-ideal interactions between organic and inorganic species exist, good agreement between measured particle water and ammonia partitioning to predictions using the bulk properties (discussed below) suggests these assumptions are reasonable.

4 Results

4.1 ~~Summary~~Overall summary of meteorology and PM composition at SOAS and SCAPE sites

For the SOAS study period, mean T and RH were 24.7 ± 3.3 °C and 73.8 ± 16.1 % (mean ± SD), respectively. This resulted in a $f(RH)$ water level of 4.52 ± 3.75 μg m^{-3}, with a maximum value of 28.41 μg m^{-3}. In comparison, SOAS mean dry $PM_{2.5}$ mass was 7.72 ± 4.61 μg m^{-3}, implying that the fine aerosols were roughly composed of 37% water, on average. Mean T and RH for SCAPE sites are listed in Table 3. Summer T means were all above 21°C, including CTR. RH means were all high (> 60%) for summer and winter, which is typical for the southeastern US.

Of the sites in the southeastern US discussed in this paper, CTR was the least influenced by anthropogenic emissions having the lowest black carbon (BC) concentrations (measured by a MAAP, Thermo Scientific, model 5012). At CTR, the mean BC = 0.26 ± 0.21 μg m^{-3} (± SD), whereas mean BC concentrations at the other rural site (YRK) was 0.36 μg m^{-3}. The representative Atlanta site (JST) BC was on average 0.71 μg m^{-3}, and higher for sites closer to roadways, 0.96 μg m^{-3} (GIT) and 1.96 μg m^{-3} (RS).

~~For the SOAS study period, mean T and RH were 24.7 ± 3.3 °C, 73.8 ± 16.1 % (mean ± SD), respectively. This resulted in $f(RH)$ water level of 4.52 ± 3.75 μg m^{-3}, with a maximum value of 28.41 μg m^{-3}. In comparison, SOAS mean dry $PM_{2.5}$ mass was 7.72 ± 4.61 μg m^{-3}, implying that the fine aerosols were composed of roughly 37% water on average. Mean T and RH for SCAPE sites are listed in Table 3. Summer T means were all above 21°C, including CTR. RH means were all high (> 60%) for summer and winter, which is typical for the southeastern US.~~

A more comprehensive suite of ions will provide a better prediction of W_i. However, in the southeastern US, inorganic ions are currently dominated by sulfate and ammonium. During SOAS, the PILS-IC provided a more comprehensive and accurate measurement of water-soluble ions than AMS, which measured only non-refractory sulfate, ammonium, nitrate, and chloride. Refractory, but water soluble ions, such as sodium and associated chloride, and crustal elements including calcium, potassium, and magnesium were present in PM_1, but in very low concentrations. Contributions of these ions are more important in $PM_{2.5}$ than for PM_1, which tend to reduce aerosol acidity. For instance, Na^+ has a significantly higher mean in $PM_{2.5}$ at 0.056 μg m^{-3} (1st half of SOAS study) than 0.001 μg m^{-3} in PM_1 (2nd half of SOAS study~~) at CTR.~~). Four, one day-long, dust events (06/12, 06/13, 06/16, and 06/21) in the SOAS data set have been excluded from this analysis as assumptions relating to internal mixing of $PM_{2.5}$

components are less valid in these cases. Excluding these days, the mean Na^+ in $PM_{2.5}$ drops to 0.024 µg m^{-3}.

If the fraction of the refractory ions (e.g., Na~~, Ca~~, $^+$, K~~, Mg~~)$^+$, Ca^{2+}, Mg^{2+}) is negligible compared to the SO_4~~, NH_4~~ (Note, SO_4 stands for sulfate in all its possible forms, from free to completely dissociated), NH_4^+, and NO_3~~,~~$^-$, the AMS data sufficiently constrains particle composition for thermodynamic calculations; this apparently is the case for most of the time in the southeast (Supplementary ~~material~~Material Section ~~5~~4). For PM_1 SO_4 and NH_4~~,~~$^+$, AMS and PILS-IC were in good agreement (~~sulfate~~SO_4 slopes within 20 %, R^2 = 0.90; ~~ammonium~~NH_4^+ within 1%, R^2 = 0.81). Similar agreement was also found for AMS PM_1 SO_4 and NH_4^+ versus PILS-IC $PM_{2.5}$ SO_4 and NH_4~~. (see~~$^+$. (See Figure 2 for comparison of complete data set). These data indicate little SO_4 and NH_4^+ between the 1.0 and 2.5 µm size range ($PM_{2.5}$ – PM_1). Because of the agreement between these dominant ions, ISORROPIA-II predicted W_i for all ions measured with the PILS-IC throughout the study (includes both PM_1 and $PM_{2.5}$) agreed with W_i based on AMS inorganic species (i.e., only ammonium and sulfate) having an orthogonal slope of 1.18, Figure 2c.

4.2 Results from the SOAS Centreville site

4.2a LWC, pH and ion balances at Centreville

The diurnal variation of LWC contributed by W_i and W_o, along with total measured water, ambient T, RH, and solar radiation at CTR is shown in Figure 3. Predicted and measured LWC trends were in good overall agreement, although the largest discrepancy was observed during the daytime when the LWC level was low and more difficult to measure and accurately predict. Nighttime RH median values were between 85% and 90% and resulted in significant water uptake that reached a peak just after sunrise near 7:30 am (local time). The dramatic peak in LWC starting at roughly 5:00 am, reaching a maximum between 7:30 and 8:00 am is likely due to RH increasing above 90%, at which point uptake of water rapidly increases with increasing RH. The similar rapid hygroscopic growth before sunrise was also observed at GIT, RS, and JST (Nov) (Figure 11). After sunrise, rising temperatures led to a rapid drop in RH, resulting in rapid loss of particle water. LWC reached lowest levels in the afternoon ~2 µg m^{-3}, only 20% of the peak value. W_o varied more than W_i diurnally; W_o max/min ratio was 13.1 compared to 4.1 for W_i.

At CTR, the aerosol was highly acidic, with predicted mean pH = 0.94 ±0.59 (±SD). The minimum and maximum pH were -0.94 and 2.23 respectively, and pH varied by approximately 1 on average throughout the day (Figure 4a). That is, the H^+_{air}/LWC ratio increased by a factor of 10 from night to day. LWC

max/min ratio was 5, whereas H^+_{air} diurnal variation was significantly less (Figure 4b), indicating that the diurnal pattern in pH was mainly driven by particle water dilution. This is further demonstrated in Figure 4d, which shows the diurnal variation in the NH_4^+/SO_4^{2-} molar ratio (the main ions driving pH), with only slightly lower ratios during the day. The study mean (±SD) NH_4^+/SO_4^{2-} molar ratio was 1.4 (±0.5). As LWC is mainly controlled by RH and temperature, the pH diurnal variation was thus largely driven by meteorological conditions, not aerosol composition.

In part, because of the diurnal variation of LWC, a simple ion balance or NH_4^+/SO_4^{2-} molar ratio or per volume air concentration of aerosol hydronium ion (H^+_{air}) alone cannot be used as a proxy for pH in the particle. Figure 5a shows a weak inverse correlation (R^2 = 0.36) between ion balance and pH. An ion balance of an aerosol is usually calculated as follows (in unit of nmol equivalence m^{-3}), for a NH_4^+-Na^+-SO_4^{2-}-NO_3^--Cl^--water inorganic aerosol.

$$Ion\ Balance = \frac{[SO_4^{2-}]}{48} + \frac{[NO_3^-]}{62} + \frac{[Cl^-]}{35.5} - \frac{[NH_4^+]}{18} - \frac{[Na^+]}{23} \quad (7)$$

Where $[SO_4^{2-}]$, $[NO_3^-]$, $[Cl^-]$, $[NH_4^+]$, and $[Na^+]$ are concentrations of these ions in units of g m^{-3}. An ion balance is also a bad indicator of pH because it poorly predicts the aerosol concentration of H^+_{air}. An ion balance assumes all ions are completely dissociated, but multiple forms are possible, depending on pH (e.g., sulfate can be in the form of H_2SO_4, HSO_4^-, or SO_4^{2-}). For example, if aerosol sulfate remains in the free form of H_2SO_4, it doesn't add protons. Thus, an ion balance usually overestimates protons and is only moderately correlated with H^+_{air} (Figure 5b).

LWC uncertainty:

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In estimating the water uncertainty, we consider W_i and W_o separately. The uncertainty of W_i is estimated by propagating the measurement uncertainty of ions and RH through the ISORROPIA-II thermodynamic model by finite perturbations about the model base state. Uncertainties of ions were estimated by difference between IC-ions and AMS-ions, as well as PILS-IC measurement uncertainty (Table 2). Na^+ is excluded because it is not measured by the AMS. PILS-IC instrumental uncertainty is estimated to be 15% from the variability in standards (variability is calibration slopes), blanks, sample airflow rate, and liquid flow rate (one ~~standard deviation).~~SD). The total ion uncertainties are listed in Table 2. SO_4 has a higher uncertainty, at 25%, than the rest, which are at 15%. These combined uncertainties lead to an W_i uncertainty of 25% (Figure ~~3), which is the same as the SO_4 uncertainty, since as the most hygroscopic ion it controls W_i uptake. Although the AMS data set does not include Na and crustal elements, it is still generally capable of predicting W_i as it measures the most hygroscopic and abundant ions SO_4, NH_4, and~~

~~NO_3 (Figure 2).~~6), which is the same as the SO_4 uncertainty. SO_4, one of the most hygroscopic ions (Petters and Kreidenweis, 2007), controls W_i uptake.

For the SOAS study, the RH probe in the ambient nephelometer (Humitter 50U, VAISALA Inc.) has a stated ~~max~~maximum uncertainty of 5% at RH = 90%. RH biases with respect to environment conditions can also occur due to placement of the probe. Based on RH comparisons between ARA, Rutgers ~~(Nguyen et al., 2014)~~(Nguyen et al., 2014), and the Georgia Tech instrumentation, a systematic bias as large as 10% is found. Given this, we consider an RH probe factory uncertainty (5%) as a typical value and inter-comparison difference (10%) as an extreme condition. In this analysis, RH was adjusted by ±5% and ±10% and W_i recalculated (Figure ~~4~~7). A ±5% perturbation in RH leads to a 91% (slope – 1~~, see Figure 7~~) error for 5% perturbation above the measured value (1.05RH) and 29% error for a perturbation below the measured value (0.95RH). We take 60% as average uncertainty. Higher uncertainty is introduced with increasing RH, owning to the exponential growth of LWC with RH and results in the asymmetric LWC uncertainty. Combining W_i uncertainty from ions (25%) and RH (60%), the overall uncertainty is calculated as 65%.

The uncertainty sources for W_o are ~~κ, ρ_s, m_s, and RH (Equation 5). The uncertainties of these parameters are estimated to be 20% (Cerully et al., 2014), 10%, 20%, and 5% (from above), respectively. In summary, the overall uncertainty of W_o is 30~~κ_{org}, ρ_s, m_s, and RH (Equation 5). The uncertainties of these parameters are estimated to be 26% (details can be found in Supplementary Material Section 3), 10%, 20%, and 5% (from above), respectively. In summary, the overall uncertainty of W_o is 35%.

The total uncertainty of LWC can be expressed as a sum of W_i and W_o uncertainties, where ~~ε_i~~ε_i is the mass fraction. ε_{W_o} was found to be 36% and ε_{W_i} was 64%.

$$\frac{\delta_{LWC}}{LWC} = \sqrt{(\varepsilon_{W_i}\frac{\delta_{W_i}}{W_i})^2 + (\varepsilon_{W_o}\frac{\delta_{W_o}}{W_o})^2}\qquad \frac{\delta_{LWC}}{LWC} = \sqrt{\left(\varepsilon_{W_i}\frac{\delta_{W_i}}{W_i}\right)^2 + \left(\varepsilon_{W_o}\frac{\delta_{W_o}}{W_o}\right)^2} \quad \text{~~(7)~~(8)}$$

Given the above, $\frac{\delta_{LWC}}{LWC}$ is 43%. This method of assessing predicted LWC uncertainty can be applied to SCAPE sites as well. The specific predicted LWC at SCAPE sites were calculated and are listed in Table 3. W_i uncertainty associated with ions is the same as noted above, 25%, because it is estimated by PILS-IC and AMS differences. Similar uncertainties in W_i at the SCAPE sites are expected if RH uncertainties are similar at all sites.

~~4.3~~ *pH uncertainty*:

As pH is based on H^+_{air} and LWC, the uncertainty of pH can be estimated from these two parameters. We applied the adjoint model of ISORROPIA, ANISORROPIA ~~(Capps et al., 2012)~~(Capps et al., 2012), to quantify the sensitivity of predicted ~~$H^\pm$~~H^+_{air} to the input aerosol species at the conditions of the thermodynamic calculations. pH uncertainty ~~resulted~~resulting from aerosol composition is then determined by propagating the input parameter uncertainties, using ANISORROPIA sensitivities, to the corresponding ~~$H^\pm$~~H^+_{air} and pH uncertainty.

We now assess how pH of $PM_{2.5}$ is affected by using an incomplete measurement of ionic species by comparing the pH predicted based on the more complete suite of ions measured by the PILS-IC versus the AMS, during SOAS. Sensitivities of aerosol species to ~~$H^\pm$~~H^+_{air} were calculated by ANISORROPIA with PILS-IC data and presented as partial derivatives (Table 2). Higher sensitivity values imply the inorganic ion is more important for ion balance. In the SOAS study, ~~$H^\pm$~~H^+_{air} is most sensitive to SO_4, and then NH_4~~$^+$~~, as they were the major ions. Uncertainties of ions were estimated by the difference between IC-ions and AMS-ions, as well as PILS-IC measurement uncertainty. Since Na^+ is not measured by AMS, we cannot estimate the difference between PILS-IC and AMS. The loadings and sensitivities of NO_3^- and Cl^- were very low, so they are assumed not to contribute much to $\frac{\delta_{H^\pm}}{H^\pm}\frac{\delta_{H^+_{air}}}{H^+_{air}}$. Given this, $\frac{\frac{\delta_{H^\pm}}{H^\pm}\delta_{H^+_{air}}}{H^+_{air}}$ is determined by;

$$\frac{\delta_{H^\pm}}{H^\pm} = \sqrt{\left(\frac{\partial H^\pm}{\partial SO_4}\frac{\delta_{SO_4}}{SO_4}\right)^2 + \left(\frac{\partial H^\pm}{\partial NH_4}\frac{\delta_{NH_4}}{NH_4}\right)^2 + \left(\frac{\partial H^\pm}{\partial Na}\frac{\delta_{Na}}{Na}\right)^2}\frac{\delta_{H^+_{air}}}{H^+_{air}}$$

$$= \sqrt{\left(\frac{\partial H^+_{air}}{\partial SO_4}\frac{\delta_{SO_4}}{SO_4}\right)^2 + \left(\frac{\partial H^+_{air}}{\partial NH_4^+}\frac{\delta_{NH_4^+}}{NH_4^+}\right)^2 + \left(\frac{\partial H^+_{air}}{\partial Na^+}\frac{\delta_{Na^+}}{Na^+}\right)^2} \quad (8)(9)$$

Based on the input for Equation ~~8~~9 (Table 2), $\frac{\delta_{H^+}}{H^+}\frac{\delta_{H^+_{air}}}{H^+_{air}}$ is estimated as 14%. LWC is most sensitive to RH fluctuations, so it is considered the main driver of LWC uncertainty in the pH calculation. As discussed, we artificially adjusted RH by ±5% and ±10% (10% is considered an extreme condition). ~~$H^\pm$~~H^+_{air}, W_i, W_o, as well as pH were all recalculated using 90%, 95%, 105%, and 110% of the actual measured RH. RH+5% and RH-5% lead to 12% and 6% variation in pH based on orthogonal regression slopes, respectively (Figure ~~5~~8). RH-10% results in only 10% variation, however, RH+10% results in a 45% variation, and the coefficient of determination (R^2) between pH calculated based on RH+10% and original RH drops to only 0.78, while for all other cases $R^2 > 0.96$. The disproportionately large effect of the positive uncertainty is caused by the exponential increase of LWC with RH, as RH reaches high levels (>90%). Assuming the

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stated manufacturer uncertainty (5%) for our RH uncertainty, pH uncertainty is estimated to be 6%-12%. We take 12% as $\frac{\partial pH}{\partial LWC}\delta_{LWC}$ for further calculations.

SO_4 was found to contribute the most to ~~$\frac{\delta_{H^\pm}}{H^\pm}$~~ $\frac{\delta_{H^+_{air}}}{H^+_{air}}$. $NH_4^\pm$ and $Na^\pm$ followed. SO_4 and $NH_4^\pm$ are the two most abundant inorganic components in aerosols and controlling aerosol acidity. Finally, the total pH uncertainty is the combination of LWC and the uncertainty associated with ~~$H^\pm$~~H^+_{air}, which is computed from the definition of pH (Equation 6).

$$\frac{\delta_{pH}}{pH} = \sqrt{\left(\frac{\partial pH}{\partial H^\pm}\delta_{H^\pm}\right)^2 + \left(\frac{\partial pH}{\partial LWC}\delta_{LWC}\right)^2}\ \frac{\delta_{pH}}{pH} = \sqrt{\left(\frac{\partial pH}{\partial H^+_{air}}\delta_{H^+_{air}}\right)^2 + \left(\frac{\partial pH}{\partial LWC}\delta_{LWC}\right)^2} \quad \text{~~(9)~~(10)}$$

where ~~$\frac{\partial pH}{\partial H^\pm}$~~ $\frac{\partial pH}{\partial H^+_{air}}$ can be derived from Equation (6) as

$$\frac{\partial pH}{\partial H^\pm}\frac{\partial pH}{\partial H^+_{air}} = -\frac{1}{2.303}\frac{1}{\frac{H^\pm}{LWC}}\frac{1}{\frac{H^+_{air}}{LWC}}\frac{1}{LWC} = -\frac{1}{2.303}\frac{1}{H^\pm}\frac{1}{H^+_{air}} \quad \text{~~(10)~~(11)}$$

From Equation 9 and the uncertainties of ~~$H^\pm$~~H^+_{air} and LWC (Equation 7 and 8), we estimate the pH uncertainty for the SOAS dataset to be 13% (based on the specific uncertainties considering here). pH uncertainties at SCAPE sites were also assessed via this method. As discussed above, ~~$\frac{\delta_{H^\pm}}{H^\pm}$~~ $\frac{\delta_{H^+_{air}}}{H^+_{air}}$ was found to be 14% for the SOAS study, due to IC and AMS data set differences and PILS-IC instrumental uncertainty. This same uncertainty is applied to SCAPE, where no PILS-IC data were available. Because aerosol ~~compositions~~composition at all sites ~~are~~is similar, based on filter IC analysis (Supplementary ~~material~~Material Figure ~~6~~S8), similar sensitivities of ~~$H^\pm$~~H^+_{air} to ions are expected. However, actual uncertainty for each sampling period is possibly higher due to higher loadings of refractory ions at SCAPE sites due to contributions from urban emissions. Refractory ions not measured by the AMS (i.e. Na^+, K^+, ~~Ca, Mg), having~~ Ca^{2+}, Mg^{2+}), have a minor effect on predicting LWC, but may have an important effect on pH (e.g., result in higher pH) in locations ~~(e.g. pH higher)~~ where they could substantially contribute to the overall ion balance.

~~***pH uncertainty in different seasons:***~~ ***4.2b Model validation: Prediction of liquid water***

Several LWC measurements were made at CTR during SOAS. In addition to $f(RH)$_water (4.52 ±3.75 μg m^{-3}), particle water was quantified with a Semi-volatile Differential Mobility Analyzer (SVDMA). With this method, a SOAS study mean particle water concentration of 4.27 ±3.69 μg m^{-3} (±STD) was obtained (Nguyen et al., 2014). The orthogonal regression between these two measurements (SVDMA water vs $f(RH)$_water) has slope = 0.91, intercept = -0.03, R^2 = 0.35. Differences could be caused by

differences in size-resolved composition (particle composition beyond PM_1 that contributes LWC; SVDMA scans up to 1.1 μm, while $f(RH)$_water is based on $PM_{2.5}$), instrument sample heating (i.e., the degree to which the instrument was close to ambient conditions, especially when ambient RH was high, and most sensitive to slight T differences), and differences in RH probe calibrations.

CTR predicted total LWC, $(W_i + W_o)$, was 5.09 ±3.76 μg m^{-3} and agreed well with $f(RH)$_water. The total predicted water was highly correlated and on average within 10% of the measured water, with slope = 0.91, intercept = 0.46, R^2 = 0.75 (see Figure 9). Since excluding refractory ions (Section 4.1) and not considering gas phase species in the ISORROPIA-II calculations do not significantly affect the LWC prediction, its comparison across sites is less uncertain than pH.

4.2c Model validation: Prediction of pH

ISORROPIA-II calculations of pH at CTR for the SOAS study were evaluated by comparing measured and predicted $NH_{3(g)}$. Although NH_4^+ and $NH_{3(g)}$, along with other aerosol components, are input into the model, comparing ambient NH_4^+ and $NH_{3(g)}$ to model predictions is not a circular analyses. For each observed data point, the model calculates total ammonia from the NH_4^+ and $NH_{3(g)}$ input, and then calculates the gas-particle ammonia partitioning assuming equilibrium. There are also other various assumptions/limitations associated with the model. Figure 10 shows the SOAS study time series of measured and predicted $NH_{3(g)}$ and the fraction of ammonia in the gas phase ($NH_{3(g)}$ / ($NH_{3(g)}$ + NH_4^+). Measured and predicted $NH_{3(g)}$ are in good agreement. Periods when almost all ammonia was in the gas phase (ratio near 1) are related to precipitation events (06/10, 06/24, 06/28, 07/03, 07/04) when aerosol concentrations were very low. Not including these events, the study mean (±SD) fraction ammonia in the gas phase was 0.41 (±0.16) (median value is also 0.41). These results provide confidence in ISORROPIA-II calculations of particle pH, and demonstrate the utility of including both measurements of particle and gas phases in these types of studies.

When gas and particle data are not available, pH predictions are not as accurate (Hennigan et al., 2014). Running ISORROPIA-II in the forward mode, but with only aerosol concentrations as input, may result in a bias in predicted pH due to repartitioning of ammonia in the model. In the southeast, where pH is largely driven by SO_4 and NH_4^+, the aerosol NH_4^+ input will be partitioned in the model between gas and particle phases to establish equilibrium. Sulfate repartitioning does not occur since it is non-volatile. Thus, NH_4^+ will be lost from the particle and a lower pH predicted. At CTR ammonia partitioning has been included in all model runs, but as no $NH_{3(g)}$ was available for SCAPE. Assuming the average $NH_{3(g)}/NH_4^+$ ratio from CTR applies to all SCAPE sites to estimate $NH_{3(g)}$, along with measured particle composition at each site, we got pH increases ranging from 0.87 to 1.38. In the following, all pHs

reported for SCAPE are corrected for this bias (i.e., pHs are increased by 1 to simplify the correction). Note that ammonia partitioning does not significantly affect the LWC prediction (W_i predicted without $NH_{3(g)}$ vs W_i predicted with $NH_{3(g)}$: slope = 1.00, intercept = -0.01 μg m^{-3}, R^2 = 0.98).

4.3 LWC and pH at other sites in the southeast (SCAPE sites)

4.3a Seasonal trends

The methods developed and verified at CTR are now applied to the SCAPE study where fewer species was measured. LWC predictions at all SCAPE sites are shown in Table 3, providing insights on seasonal trends of LWC in the southeast. The overall summer LWC mean was 5.02 μg m^{-3} and winter mean 2.22 μg m^{-3}.

At the SCAPE sites, JST, YRK, GIT, and RS, summer mean pHs were between 1 and 1.3, similar to CTR (mean of 0.94). In winter the pHs (mean between 1.8 and 2.2) were higher by ~ 1 unit. Although LWC was higher in summer, which tends to dilute H^+_{air} and increase pH, summer pH was lower due to higher ion (i.e., sulfate) concentrations (Table 3). Similar diurnal pH patterns were seen at all sites in all seasons and follow the diurnal variations of particle water (Figure 11). Overall the pH in the southeast is very low, between 1 and 2 (mean), in both rural and urban environments. pH values in summer at various sites were similar (1 to 1.3), suggesting a fairly homogeneous distribution of acidity due to spatially uniform sulfate in the southeastern US (Zhang et al., 2012). In winter the diurnal range in pH was roughly 2 units, while the diurnal range in summer was smaller, with pH varying by roughly 1.

Recall at CTR, 10% RH uncertainty can result in a pH prediction error of up to 45% due to the high RHs observed during the study. We estimated pH uncertainty from W_i and W_o by + 10% RH for each SCAPE site. As Table 3 shows, the pH uncertainty associated with RH is much lower in winter (only 1-3%) than summer (20-40%), although RH averages were similar, e.g., JST in May (67 ±19%) and Nov (63 ±19%), with even higher RH in winter at YRK. Total pH uncertainty at all SCAPE sites are calculated by the same method as CTR. Table 3 shows that higher RH and T result in larger pH uncertainty. In summer, pH uncertainty is mainly caused by RH; while in winter, it can be attributed mostly to uncertainty in ion concentrations. ~~pH predicted for the SCAPE sites likely has a systematic bias of roughly -1 because $NH_{3(g)}$ is not included in the ISORROPIA-II calculation (recall including $NH_{3(g)}$ increased the CTR pH by 0.8). At JST, YRK, GIT, and RS sites, summer mean pH were between 0 and 0.3. Adding 1 pH unit of systematic bias makes pH at these southeastern sites between 1 and 1.3, which is consistent with the CTR pH of 0.94 ±0.59 (±SD). $NH_{3(g)}$ does not significantly affect the LWC prediction (W_t predicted without $NH_{3(g)}$ vs W_t predicted with $NH_{3(g)}$: slope = 1.00, intercept = -0.01 μg m^{-3}, R^2 = 0.98).~~

578 **4.4 LWC comparison to predictions and seasonal and diurnal trends in the southeast**

579 ***4.3b The role of W_o***

580 Several LWC measurements were made at CTR during SOAS. In addition to $f(RH)$_water (4.52 ±3.75
581 μg m^{-3}), particle water was quantified with a Semi-volatile Differential Mobility Analyzer (SVDMA).
582 With this method, a SOAS study mean particle water concentration of 4.27 ±3.69 μg m^{-3} (±SD) was
583 obtained (Nguyen et al., 2014). The orthogonal regression between these two measurements (SVDMA
584 water vs $f(RH)$_water) has slope = 0.91, intercept = −0.03, r = 0.59. Differences could be caused by
585 differences in size-resolved composition (particle composition beyond PM_1 that contributes LWC;
586 SVDMA scans up to 1.1 μm, while $f(RH)$_water is based on $PM_{2.5}$), instrument sample heating (i.e., the
587 degree to which the instrument was close to ambient conditions, especially when ambient RH was high,
588 hence most sensitive to slight T differences), and differences in RH probe calibrations.

589 The particle water predicted from the sum of W_t and W_o agreed well with $f(RH)$_water. That is, the total
590 predicted water was highly correlated and on average within 10% of the measured water, with slope =
591 0.91, intercept = 0.46, r = 0.87 (see Figure 6).

592 Since excluding refractory ions (Section 4.1) and not considering gas phase species in the ISORROPIA-II
593 calculations do not significantly affect the LWC prediction, its comparison across sites is less uncertain
594 than pH. CTR total LWC is predicted to be 5.09 ±3.76 μg m^{-3}. LWC predictions at all the other sites are
595 shown in Table 3, providing insights on seasonal trends of LWC in the southeast. The overall summer
596 LWC mean was 5.02 μg m^{-3} and winter mean 2.22 μg m^{-3}.

597 The diurnal variation of LWC contributed by W_t and W_o, along with total measured water, ambient T, RH,
598 and solar radiation at CTR is shown in Figure 7. Predicted and measured LWC trends are in good overall
599 agreement, although the largest discrepancy was observed during the daytime when LWC level was low
600 and more difficult to measure and accurately predict. Nighttime RH median values were between 85%
601 and 90% and resulted in significant water uptake that reached a peak just after sunrise near 7:30 am (local
602 time). The dramatic peak in LWC starting at roughly 5:00 am, reaching a maximum between 7:30 and
603 8:00 am is likely due to RH increasing above 90%, at which point uptake of water rapidly increases with
604 increasing RH. The similar rapid hygroscopic growths before sunrise were also observed at GIT, RS, and
605 JST (Nov). After sunrise, rising temperatures lead to a rapid drop in RH, resulting in rapid loss of particle
606 water. LWC reached lowest levels in the afternoon ~2 μg m^{-3}, only 20% of the peak value. W_o varied
607 more than W_t; W_o diurnal variation ratio (max/min) was 13.1 compared to 4.1 for W_t.

608 W_o was significant, accounting for on average 29-39% of the total $PM_{2.5}$ particle water for all our sites
609 (Figure ~~8~~12 and Table 3). Note that, W_o at SCAPE sites were calculated by in-situ AMS
610 ~~measurement~~measurements at each SCAPE site and ~~measured~~the mean κ_{org} ~~mean~~ (0.126) measured at
611 CTR, due to lack of CCNc. Note that ε_{W_o} could be higher or lower at each site depending on the type of
612 organics presented and the related κ_{org}. Figure ~~8~~12 shows that W_o is related to the organic mass fraction.
613 W_o is comparable to W_i at night. In contrast, it was only 33% of W_i during the daytime (Figure ~~7~~3). The
614 significant fraction, even during daytime, indicates organic aerosol components will have a considerable
615 contribution to aerosol radiative forcing. Although organics are less hygroscopic than ammonium sulfate,
616 a large fraction of the $PM_{2.5}$ (~70%) was organic, making W_o contributions important. Of the organic
617 factors associated with W_o, ~~Cerully et al. (2014)~~Cerully et al. (2014) showed that MO-OOA (more-
618 oxidized oxygenated organic aerosol, also referred to as LVOOA, low-volatile oxygenated organic
619 aerosol) and Isoprene-OA (isoprene derived organic aerosol) were twice as hygroscopic as LO-OOA
620 (less-oxidized oxygenated organic aerosol, also referred to as SVOOA, semi-volatile oxygenated organic
621 aerosol). The LWC associated with MO-OOA and Isoprene-OA account for ~60% and ~30% of total W_o
622 in the daytime, respectively.

623 **~~4.5 Ammonia partitioning, model validation and pH bias when gas/particle data not included~~**

624 ~~As noted earlier, ISORROPIA-II calculations at CTR for the SOAS study were evaluated by comparing~~
625 ~~measured and predicted $NH_{3(g)}$. Although NH_4 and $NH_{3(g)}$, along with other aerosol components, are input~~
626 ~~into the model, comparing ambient NH_4 and $NH_{3(g)}$ to model predictions is not a circular analyses. For~~
627 ~~each observed data point, the model calculates total ammonia from the NH_4 and $NH_{3(g)}$ input, and then~~
628 ~~calculates the gas-particle ammonia partitioning assuming equilibrium and the other various~~
629 ~~assumptions/limitations associated with the model. Figure 9 shows the SOAS study time series of~~
630 ~~measured and predicted $NH_{3(g)}$ and the fraction of ammonia in the gas phase ($NH_{3(g)}$ / ($NH_{3(g)}$ + $NH_{4(p)}$).~~
631 ~~Diurnal changes in the gas fraction can be seen and are likely due to diurnal changes in particle pH,~~
632 ~~discussed below.~~The effect of aerosol sources of particle water on pH can also be delineated. pH
633 calculated just by W_i alone will be affected by an underestimation of particle water, resulting in a slightly
634 lower pH (Figure 13). W_o is on average 29% to 39% of total water at all sites, as a result pH increases by
635 0.15 to 0.23 units when W_o is included. Independent of the pH range, a 29% to 39% W_o fraction always
636 increases pH by 0.15 to 0.23 due to the logarithmic nature of pH. The effect of W_o on pH can be simply
637 denoted as $\log_{10}(1-\varepsilon_{W_o})$. For example, when ε_{W_o} is 90%, it shifts pH up by 1 unit. pH based on W_i is
638 highly correlated with pH for total water ($W_i + W_o$) (Slope = 0.94, intercept = -0.14, R^2 = 0.97). This
639 indicates that if organic mass and κ_{org} are not available, ISORROPIA-II run with only ion data will give

a reasonable estimate of pH, since both H^+_{air} and W_i are outputs of ISORROPIA-II, while W_o is predicted based on organic mass and κ_{org}. Accurate temperature and RH are still necessary inputs, especially when RH is high.

4.4 Overall implications of low pH

Highly acidic aerosols throughout the southeast during all seasons~~. Periods when almost all ammonia was in the gas phase (ratio near 1) are related to precipitation events (06/10, 06/24, 06/28, 07/03, 07/04) when aerosol concentrations were very low. Not including these events, the study mean (±SD) fraction ammonia in the gas phase was 0.41 (±0.16) (median value is also 0.41). These results provide confidence in ISORROPIA-II calculations of particle pH, and demonstrate the utility of including both measurements of particle and gas phases in these types of studies.~~

~~When gas and particle data are not available, pH predictions are likely not as accurate (Hennigan et al., 2014). Running ISORROPIA-II in the forward mode, but with only aerosol concentrations as input, may result in a bias in predicted pH due to repartitioning of ammonia in the model. In the southeast, where pH is largely driven by SO_4 and NH_4, the aerosol NH_4 input will be partitioned in the model between gas and particle phases to establish equilibrium. Sulfate repartitioning does not occur since it is non-volatile. Thus NH_4 will be lost from the particle and a lower pH predicted. At CTR ammonia partitioning has been included in all model runs, but as no $NH_{3(g)}$ was available for SCAPE, the pH is likely biased at these sites by roughly -1. In the following, all pHs reported for SCAPE are corrected for this bias (i.e., pH increased by 1). Note that ammonia partitioning does not significantly affect the LWC prediction (W_t predicted without $NH_{3(g)}$ vs W_t predicted with $NH_{3(g)}$; slope = 1.00, intercept = -0.01 μg m^{-3}, R^2 = 0.98).~~

~~4.6 pH diurnal, seasonal, and spatial trends in the southeast~~

~~At CTR, the aerosol was highly acidic, with predicted mean pH = 0.94 ± 0.59 (±SD). The minimum and maximum pH were -0.94 and 2.23 respectively, and pH varied by approximately 1 on average throughout the day (Figure 10). That is, the H^+_{air}/LWC ratio increased by a factor of 10 from night to day. LWC max/min ratio was 5, whereas H^+ diurnal variation was significantly less (Figure 10), indicating that the diurnal pattern in pH was driven mainly by particle water dilution. This is further demonstrated in Figure 11, which shows the diurnal variation in the NH_4/SO_4 molar ratio (the main ions driving pH), with only slightly lower ratios during the day. The study mean (±SD) NH_4/SO_4 molar ratio was 1.4 (±0.5). As LWC is mainly controlled by RH and temperature, the pH diurnal variation was thus largely driven by meteorological conditions, not aerosol composition. Because of this, a simple ion balance or per volume air concentration of aerosol hydronium ion (H^+_{air}) alone cannot be used as a proxy for pH in the particle.~~

~~At the SCAPE sites, JST, YRK, GIT, and RS, summer mean pHs were between 1 and 1.3, similar to CTR (mean of 0.94). In winter the pHs (mean between 1.8 and 2.2) were higher by ~ 1 unit. Although LWC was higher in summer, which tends to dilute H^+ and increase pH, summer pH was lower due to higher ion (i.e., sulfate) concentrations (Table 3). Similar diurnal pH patterns were seen at all sites in all seasons and follow the diurnal variations of particle water (Figure 12). Overall the pH in the southeast is very low, between 1 and 2 (mean), in both rural and urban environments. pH values in summer at various sites were similar (1 to 1.3), suggesting a fairly homogeneous distribution of acidity due to spatially uniform sulfate in the southeastern US (Zhang et al., 2012). In winter the diurnal range in pH was roughly 2, while the diurnal range in summer was smaller, with pH varying by roughly 1.~~

~~These acidic aerosols in the southeast~~ will affect a variety of processes. For example, aerosol acidity strongly shifts the partitioning of $HNO_{3(g)}$ to the gas phase resulting in low nitrate aerosol levels in the southeast during summer (the higher summertime temperature also plays a secondary role). Aerosol acidity also impacts the gas-particle partitioning of semi-volatile organic acids. ~~(~~Note, organic acids are not considered in our model, under these acidic conditions (pH = 1) their contributions to the ~~ion balance~~ H^+_{air} (hence pH) are expected to be negligible. Because the pK_a (pK_a = - $\log_{10} K_a$, K_a referred as acid dissociation constant) of trace organic acids are > 2 (e.g~~.~~., pK_a of formic acid, one of the strongest organic acids, is 3.75, Bacarella et al. (1955)), low pH prevents dissociation of the organic acids. Since H^+ is involved in aqueous phase reactions, low pH can affect reaction rates by providing ~~more~~ protons. Investigators have found that ~~isoprene~~Isoprene-OA formation is acid-catalyzed and sulfuric acid participates in the reaction as a proton donor in chamber studies ~~(Surratt et al., 2007)~~(Surratt et al., 2007). However, aerosol acidity ~~may~~appears not to be ~~the~~a limiting factor ~~of isoprene~~for Isoprene-OA formation in the southeastern US, owning to the consistently very low pH ~~(Karambelas et al., 2014; Xu et al., 2014)~~(Karambelas et al., 2014; Xu et al., 2015). Finally, low pH can affect the solubility of trace metals ~~found in the insoluble fraction of aerosol~~ (e.g., mineral dust) such as Fe and Cu, which possibly increases the toxicity of the redox metals ~~(Ghio et al., 2012; Verma et al., 2014)~~(Ghio et al., 2012; Verma et al., 2014) and may also have a long term effect on nutrient distributions in the region (~~Meskhidze et al., 2003; Meskhidze et al., 2005~~Meskhidze et al., 2003; Meskhidze et al., 2005; ~~Nenes et al., 2011~~Nenes et al., 2011; Ito and Xu, 2014).

Field Code Changed

~~Finally, the effect of aerosol sources of particle water on pH can also be delineated. pH calculated just by W_t alone will be affected by an underestimation of particle water, resulting in a slightly lower pH (Figure 13). W_o is on average 29% to 39% of total water at all sites, as a result pH increases by 0.15 to 0.23 units when W_o is included. pH based on W_t is highly correlated with pH for total water (W_t + W_o) (R^2 = 0.97, Figure 14). This indicates that if organic mass and κ_{org} are not available, ISORROPIA-II run with only~~

~~ion data will give a reasonable estimate of pH, since both H^+_{air} and W_t are outputs of ISORROPIA-II, while W_o is predicted based on organic mass and κ_{org}. Accurate temperature and RH are still necessary inputs, especially when RH is high.~~

5 Conclusions

~~Because particle pH is important and difficult to measure directly, by combining several models we present a comprehensive prediction method to calculate pH and include an uncertainty analysis. ISORROPIA-II is applied to calculate the concentration of H^+ and W_t from inorganic aerosol measurement, and CCN activity is used to predict W_o. The adjoint model of ISORROPIA, ANISORROPIA, is applied to determine sensitivities, which are used for propagating the measurement uncertainties to pH.~~

~~W_o~~ Particle pH is important and difficult to measure directly. However, the commonly used pH proxies of ion balances and NH_4^+/SO_4^{2-} molar ratios don't necessarily correlate with pH. Therefore, predicting pH is the best method to analyze particle acidity. By combining several models we present a comprehensive prediction method to calculate pH and include an uncertainty analysis. ISORROPIA-II is applied to calculate the concentration of H^+_{air} and W_i from inorganic aerosol measurements, and CCN activity is used to predict W_o. The adjoint model of ISORROPIA, ANISORROPIA, is applied to determine sensitivities, which are used for propagating the measurement uncertainties to pH. We find that W_o should be included when predicting particle LWC when organic loadings are high (such as in the southeastern US). However, the pH prediction is not highly sensitive to W_o, unless W_o mass fraction to the total particle water is close to 1. Thus, in most cases particle pH can be predicted fairly accurately with just measurements of inorganic species and ISORROPIA-II. However, constraining ISORROPIA-II with gas phase species, such as $NH_{3(g)}$, as done in this work (or $HNO_{3(g)}$), is highly recommended, along with running ISORROPIA-II in the forward mode. ISORROPIA-II does not consider organic acids, but at the low pHs of this study, they do not contribute protons (Bacarella et al., 1955). However, for pH approaches 7, the dissociation of organic acids cannot be neglected. Finally, the model was validated through comparing predicted to measured liquid water (W_i+W_o to $f(RH)$_water) and predicted to measured $NH_{3(g)}$ concentrations.

On average, for the SOAS and SCAPE field studies, particle water associated with the $PM_{2.5}$ organic species (W_o) accounted for a significant fraction of total LWC, ~~at~~with a mean of 35% (±3% SD) ~~for SOAS and SCAPE field studies,~~ indicating the importance of organic hygroscopic properties to ~~aerosol scattering~~aqueous phase chemistry and ~~direct~~ radiative forcing in the southeast US. Although organics are

less hygroscopic than sulfate and ammonium, the larger mass fraction of organics than inorganics promotes W_o uptake. Predicted LWC was compared to LWC determined from ambient versus dry light scattering coefficients and a TEOM ~~measuring~~ measurement of dry $PM_{2.5}$ mass. In SOAS, ~~measured LWC~~ the sum of W_i and W_o was highly correlated and in close agreement with the ~~sum of W_i and W_o~~ measured LWC (slope = 0.91, ~~r~~R^2 = 0.~~87~~75). LWC showed a clear diurnal pattern, with a continuous increase at night (median ~~nighttime LWC~~ of 10 µg m^{-3} at 7:30 am) reaching a distinct peak when RH reached a maximum near 90% just after sunrise ~~at~~during the period of lowest ~~ambient~~daily temperature, followed by a rapid decrease and lower values during the day (median of 2 µg m^{-3} at 2:30 pm).

In the southeastern US, pH normally ~~varies~~varied from 0.5 to 2 in the summer and 1 to 3 in the winter, indicating that the aerosol ~~is~~was highly acidic throughout the year. The minimum and maximum pH were -0.94 and 2.2 at CTR, respectively and varied from a nighttime average of 1.5 ~~at night~~ to daytime average of 0.6 ~~on average in the day~~, mostly attributable to diurnal variation in RH and temperature. Mean ~~ammonium to sulfate~~NH_4^+/SO_4^{2-} molar ratios were 1.4 ±0.5 (SD) and roughly half the ammonia was in the gas phase ($NH_{3(g)}$ / ($NH_{3(g)}$ + $NH_{4(p)}^+$) = 41 ±16 %, mean ±SD). pH ~~throughout~~at other sites in the southeast (~~CTR excluded~~SCAPE study) was estimated based on a limited data set at an estimated uncertainty of 9-49% and a systematic bias of -1 ~~if~~since $NH_{3(g)}$ is not included in the thermodynamic model run in the forward mode. pH can still be predicted with only aerosol measurements, but ~~a~~an adjustment of 1-unit pH increase ~~adjustment~~ is recommended for the southeastern US. pH has a diurnal trend that follows LWC, higher (less acidic) at night and lower (more acidic) during the day. pH was also generally higher in the winter (~2) than summer (~1). These low pHs have significant implications ~~on~~for gas-aerosol partitioning, acid-catalyzed reactions including isoprene-OA formation, and trace metal mobilization.

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1054 **Table. 1.** Deployment status of instruments at various sites. All the listed instruments or probes were
1055 operated at CTR for SOAS.

Site	Period (mm yyyy)	PILS-IC	AMS	CCNc	Nephelometer	TEOM	RH&T
JST	May&Nov 2012	NO	YES	NO	NO	YES	YES
YRK	Jul&Dec 2012	NO	YES	NO	NO	YES	YES
GIT	Jul-Aug 2012	NO	YES	NO	NO	YES	YES
RS	~~Sep~~Sept 2012	NO	YES	NO	NO	YES	YES
CTR	Jun-Jul 2013	YES	YES	YES	YES	YES	YES

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Table. 2. Sensitivity of $~~H^+~~H^+_{air}$ to ions from ANISORROPIA (2nd row) and contribution ~~of~~to uncertainty. Uncertainties of inorganic ions ($\frac{\delta_{Ion}}{Ion}$) are calculated based on a combination of PILS-IC instrumental relative uncertainties (IC uncertainty, referred to as $\frac{\delta_{ion,IC}}{Ion}$, all estimated to be 15%) and the difference between PILS-IC and AMS ($\frac{\delta_{Ion,IC-AMS}}{Ion}$, defined as the (slope – 1) in Figure ~~1~~2a & 2b) (3rd row), where $\frac{\delta_{Ion}}{Ion} = ~~\sqrt{\left(\frac{\delta_{ion,IC}}{Ion}\right)^2 + \left(\frac{\delta_{Ion,IC-AMS}}{Ion}\right)^2}~~ \sqrt{\left(\frac{\delta_{ion,IC}}{Ion}\right)^2 + \left(\frac{\delta_{Ion,IC-AMS}}{Ion}\right)^2}$ (4th row). Contribution of uncertainty is the ratio of ion uncertainty over $~~H^+~~H^+_{air}$ uncertainty ($\frac{\delta_{~~H^+~~}}{~~H^+~~}\frac{\delta_{H^+_{air}}}{H^+_{air}}$, calculated to be 14% in Equation 8) (5th row).

PILS-IC ion concentration, μg m^{-3} (mean ± SD)	SO_4	NH_4^+	Na^+	NO_3^-	Cl^-
	1.73 ±1.21	0.46 ±0.34	0.03±0.07	0.08 ±0.08	0.02 ±0.03
$~~H^+~~H^+_{air}$ Sensitivity (mean ± SD)	$\left\|\frac{\partial ~~H^+~~}{\partial ~~SO_4~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial SO_4}\right\|$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~NH_4~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial NH_4^+}\right\|$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~Na~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial Na^+}\right\|$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~NO_3~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial NO_3^-}\right\|$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~Cl~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial Cl^-}\right\|$
	0.51 ±0.34	0.32 ±0.31	0.19 ±0.27	0.002 ±0.007	0.000 ±0
$\frac{\delta_{Ion,IC-AM}}{Ion}$	$\frac{\delta_{SO_4,IC-AMS}}{SO_4}$	$\frac{~~\delta_{NH_4,IC-AMS}~~}{~~NH_4~~}\frac{\delta_{NH_4^+,IC}}{NH}$	$\frac{~~\delta_{Na,IC-AMS}~~}{~~Na~~}\frac{\delta_{Na^+,I}}{N}$	$\frac{~~\delta_{NO_3,IC-AMS}~~}{~~NO_3~~}\frac{\delta_{NO_3^-,IC}}{NO_3}$	$\frac{~~\delta_{Cl,IC-AMS}~~}{~~Cl~~}\frac{\delta_{Cl^-,IC}}{Cl}$
	20.5%	1.5%	N/A*	**	**
$\frac{\delta_{Ion}}{Ion}$	$\frac{\delta_{SO_4}}{SO_4}$	$\frac{~~\delta_{NH_4}~~}{~~NH_4~~}\frac{\delta_{NH_4^+}}{NH_4^+}$	$\frac{~~\delta_{Na}~~}{~~Na~~}\frac{\delta_{Na^+}}{Na^+}$	$\frac{~~\delta_{NO_3}~~}{~~NO_3~~}\frac{\delta_{NO_3^-}}{NO_3^-}$	$\frac{~~\delta_{Cl}~~}{~~Cl~~}\frac{\delta_{Cl^-}}{Cl^-}$
	25.4%	15.1%	15%	15%	15%
Contribution to $~~H^+~~H^+_{air}$ uncertainty	$\left\|\frac{\partial ~~H^+~~}{\partial ~~SO_4~~}\right\|\cdot\frac{~~\delta_{SO_4}~~}{~~SO_4~~}\left\|\frac{\partial H}{\partial S}\right\|\frac{}{~~\delta_{H^+}~~}$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~NH_4~~}\right\|\cdot\frac{~~\delta_{NH_4}~~}{~~NH_4~~}\left\|\frac{\partial H}{\partial N}\right\|\frac{}{~~\delta_{H^+}~~}$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~Na~~}\right\|\cdot\frac{~~\delta_{Na}~~}{~~Na~~}\left\|\frac{\partial H^+}{\partial Na}\right\|\frac{}{~~\delta_{H^+}~~}$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~NO_3~~}\right\|\cdot\frac{~~\delta_{NO_3}~~}{~~NO_3~~}\left\|\frac{\partial H}{\partial N}\right\|\frac{}{~~\delta_{H^+}~~}$	$\left\|\frac{\partial ~~H^+~~}{\partial ~~Cl~~}\right\|\cdot\frac{~~\delta_{Cl}~~}{~~Cl~~}\left\|\frac{\partial H}{\partial C}\right\|\frac{}{~~\delta_{H^+}~~}$
	0.93	0.35	0.20	0.002	0.000

* Na^+ is not measured by AMS.

** $\left\|\frac{\partial ~~H^+~~}{\partial ~~NO_3~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial NO_3^-}\right\|$ and $\left\|\frac{\partial ~~H^+~~}{\partial ~~Cl~~}\right\|\left\|\frac{\partial H^+_{air}}{\partial Cl^-}\right\|$ are less than 1% of the other $~~H^+~~H^+_{air}$ sensitivities, and the loadings of NO_3^- and Cl^- are less than 5% of the total inorganic ion mass. As a result, their contributions to $~~H^+~~H^+_{air}$ uncertainty are negligible.

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~~Particle water and~~Fine particle pH in the southeastern United States

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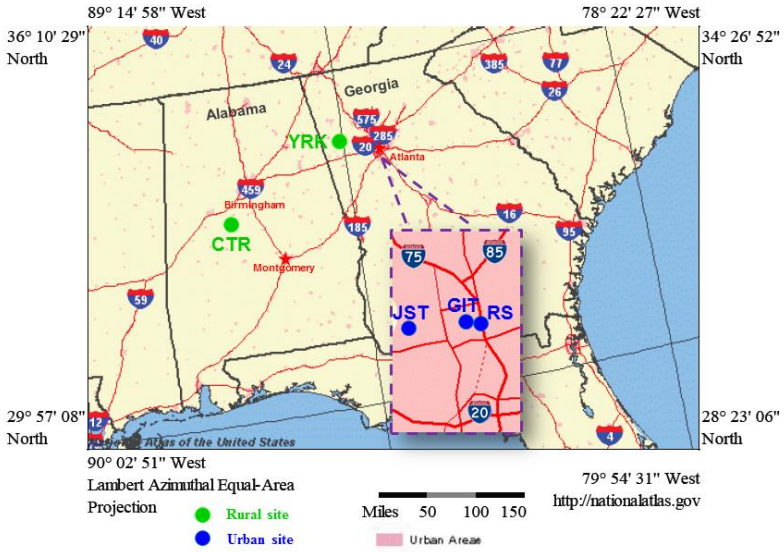
Table. 3. Water and pH prediction for SCAPE sites. Means and SDs are listed, if not specified. Total ion concentration is counted as the sum of AMS inorganics (3rd row). ~~$\epsilon(W_o)$~~ε_{W_o} is the mass fraction of W_o (5th row).

	JST 05/2012	YRK 07/2012	GIT 08/2012	RS 09/2012	JST 11/2012	YRK 12/2012
RH, %	67 ±19	66 ±21	71 ±17	72 ±20	63 ±19	73 ±21
T, ℃	23.1 ±4.3	27.7 ±4.4	26.3 ±3.5	21.4 ±3.8	11.5 ±4.8	9.8 ±5.2
Total ion concentration, μg m^{-3}	4.1 ±2.1	4.5 ±2.2	5.3 ±2.6	4.1 ±2.7	3.6 ±2.1	2.3 ±1.8
$\frac{\delta_{pH}}{pH}$ from 1.10RH	22.3%	21.4%	48.3%	22.1%	2.5%	1.4%
Total $\frac{\delta_{pH}}{pH}$	23.9%	23.0%	49.0%	23.7%	8.8%	8.6%
~~$\epsilon(W_o)$~~, ε_{W_o}, %	34 ±11	37 ±8	33 ±10	38 ±11	39 ±16	29 ±15
LWC, μg m^{-3}	5.98 ±6.28	8.14 ±8.47	8.41 ±7.67	7.81 ±9.23	5.88 ±8.69	3.24 ±3.46
pH*	1.3 ±0.7	1.1 ±0.6	1.1 ±0.4	1.3 ±0.7	2.2 ±0.9	1.8 ±1.0
LWC, μg m^{-3} (median)	3.74 ±6.28	5.29 ±8.47	6.06 ±7.67	4.31 ±9.23	2.14 ±8.69	2.02 ±3.46
pH* (median)	1.2 ±0.7	1.0 ±0.6	1.0 ±0.4	1.2 ±0.7	2.3 ±0.9	1.8 ±1.0

* A bias correction of 1 pH unit is applied due to not considering ammonia partitioning. See Section 4.2c for details.



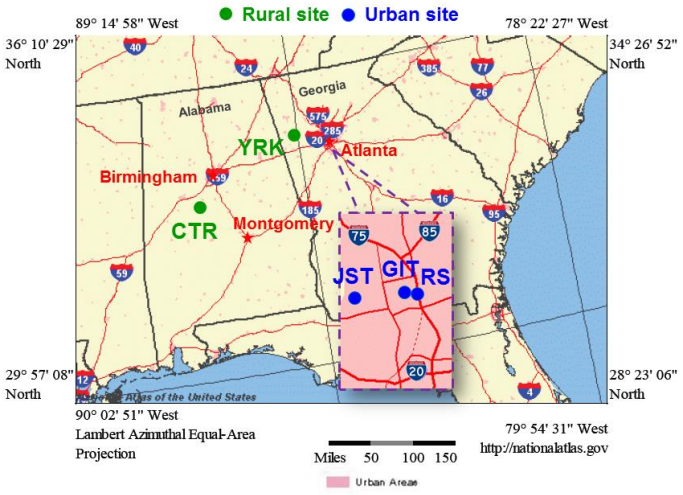


Fig. 1. Sampling sites in the southeastern ~~USA~~US, consisting of two rural and three urban sites.

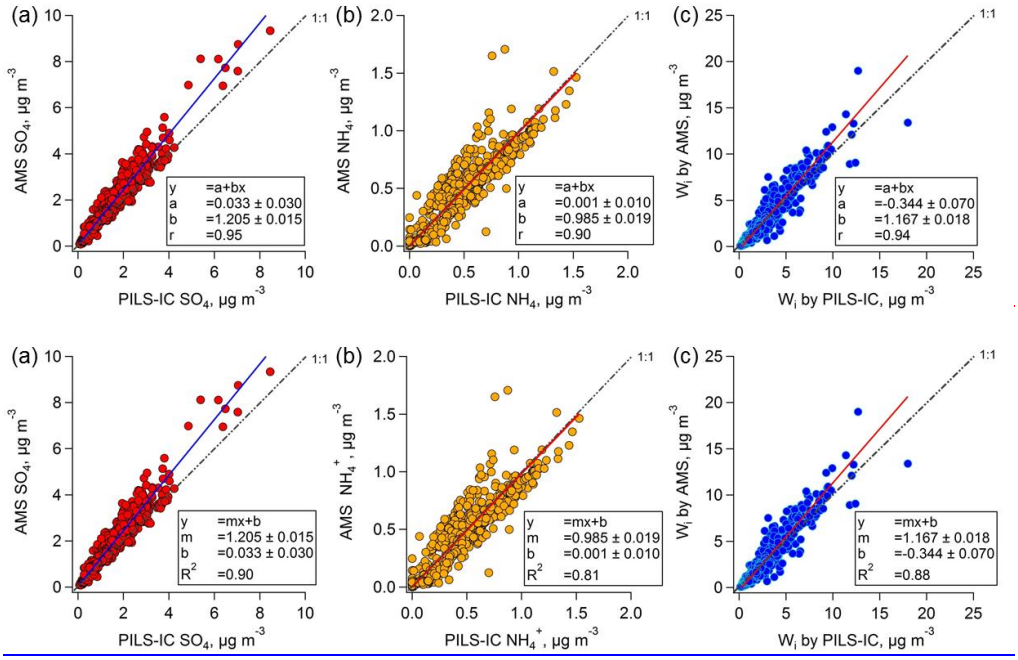


Fig. 2. Comparisons of PM_1 AMS sulfate, ammonium to PM_1 and $PM_{2.5}$ PILS-IC (i.e. complete SOAS study) and predicted W_i. Orthogonal distance regression (ODR) fits were applied.

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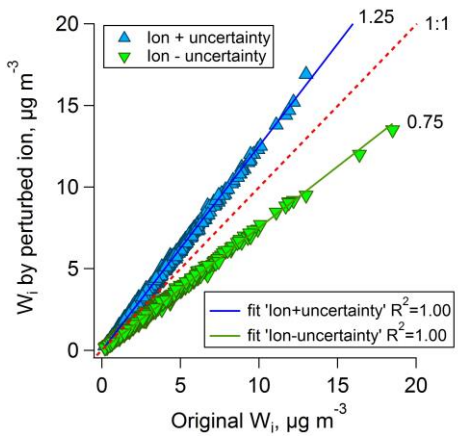

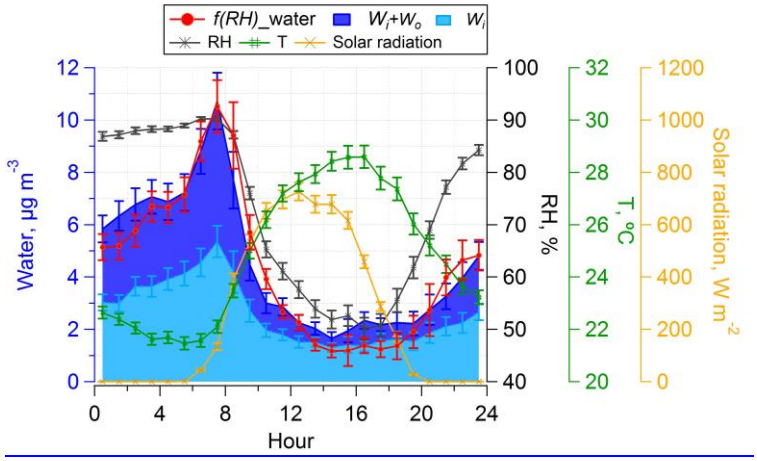

Fig. 3. CTR (SOAS) diurnal profiles of predicted and measured water, measured RH, T, and solar radiation. Median hourly averages are shown and standard errors are plotted as error bars.

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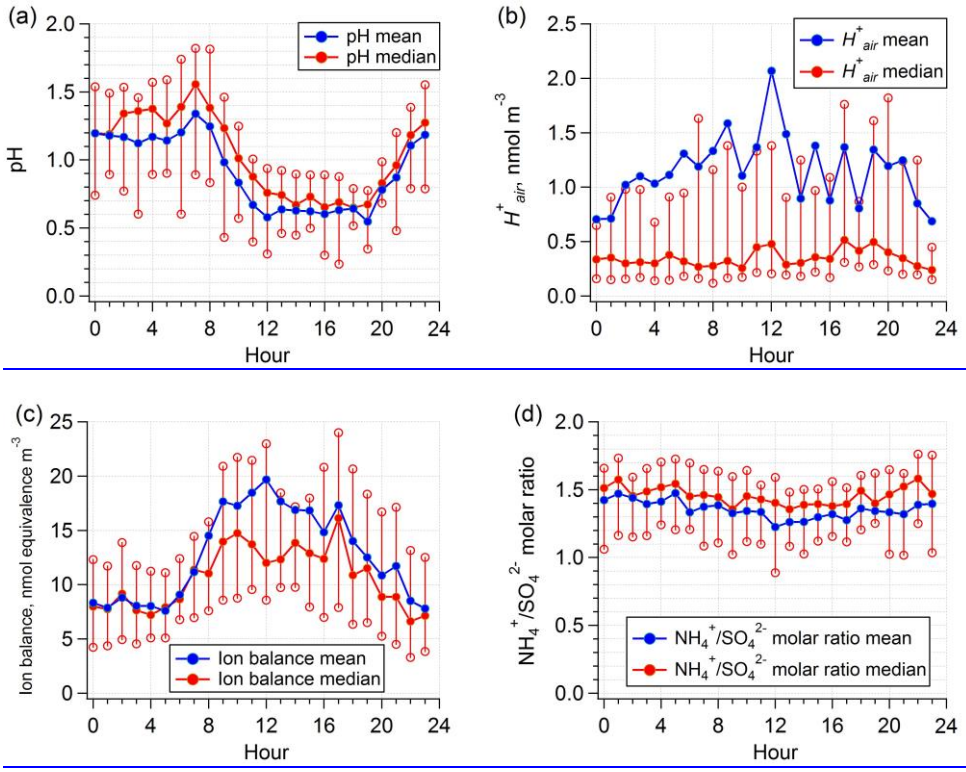


Fig. ~~W_t~~4. CTR (SOAS) diurnal patterns of calculated pH based on ~~artificially perturbed~~total predicted water (W_i + W_o) (a), H^+_{air} predicted by ISORROPIA-II (b), ion balance (c), and NH_4^+/SO_4^{2-} molar ratio (d). Mean and median values are shown, together with 25% and 75% quantiles marked as non-filled circles.

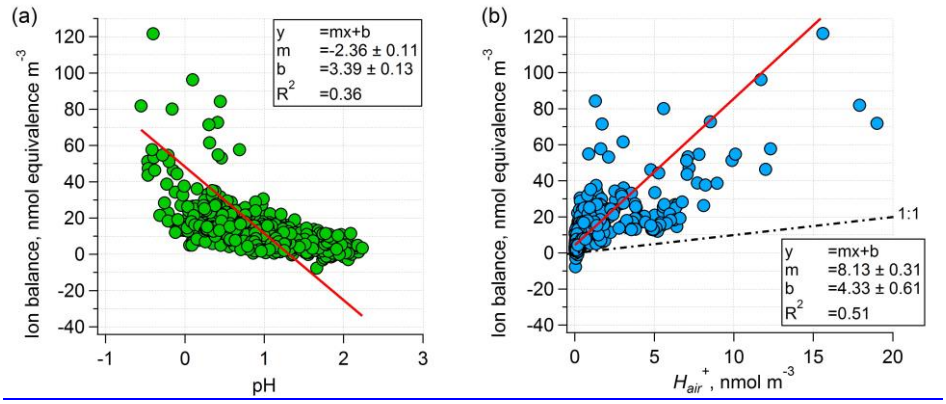


Fig. 5. Comparison of ion balance to pH (a) and to H^+_{air} (b) at CTR (SOAS). An ODR fit was applied.

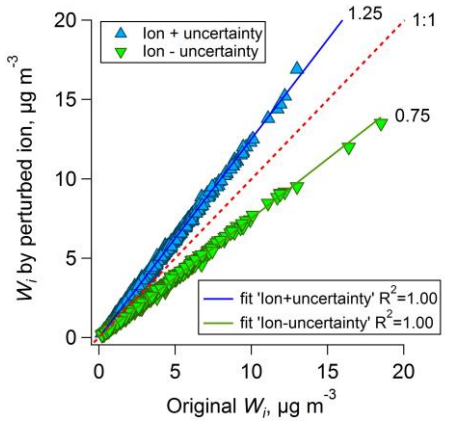


Fig. 6. W_i based on artificially perturbed ion data at upper and lower uncertainty limits~~-~~ is compared to W_i at base level. The slopes indicate the W_i uncertainty caused by ions.

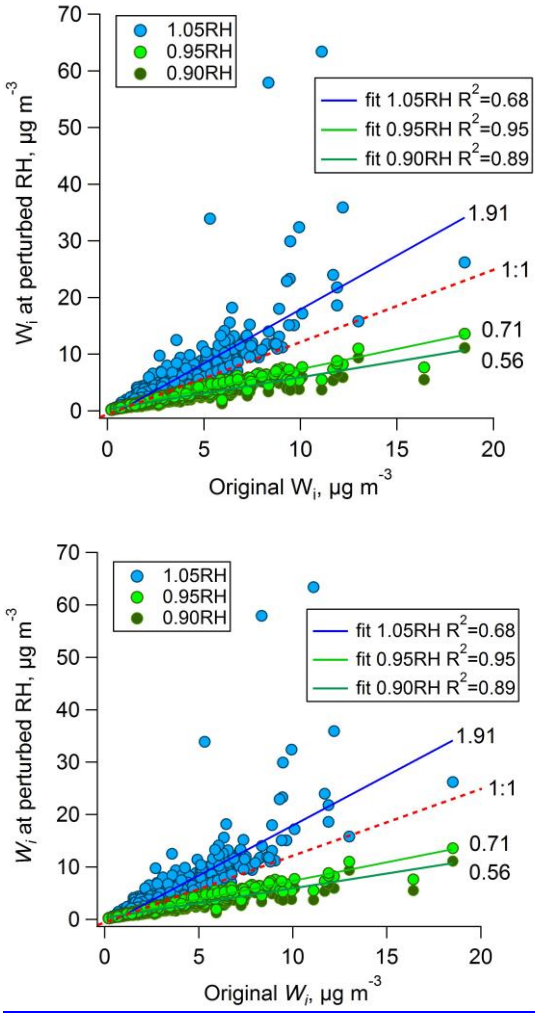
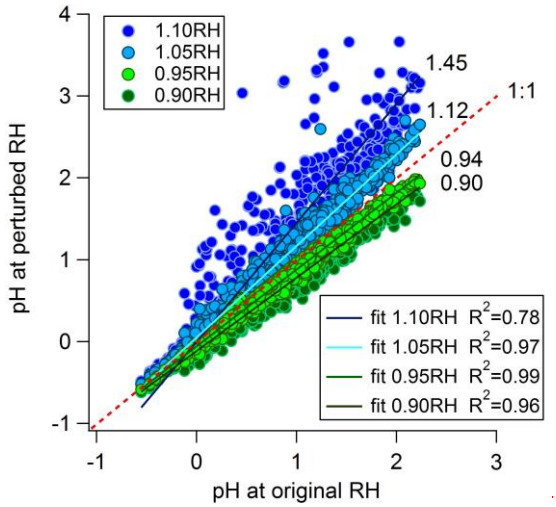


Fig. ~~4~~7. W_i based on artificially perturbed RH at upper and lower uncertainty limits compared to W_i at base level. 1.10RH (i.e., RH increased by 10%) is not plotted because it results in much larger W_i than the rest. Slopes and R^2 indicate corresponding W_i uncertainty caused by variability (uncertainty) in RH.



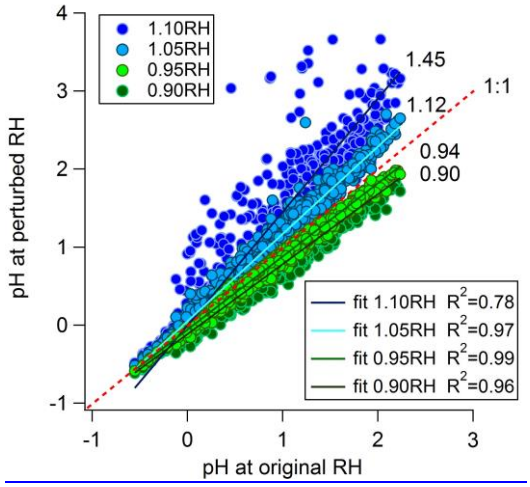


Fig. ~~5~~8. pH predictions by ~~perturbed~~perturbing RH compared to pH at base level. W_i, W_o, and H^+_{air} were recalculated based on ±5% and ±10% original RH to investigate pH uncertainty. The slopes and R^2 indicate pH uncertainty caused by RH.

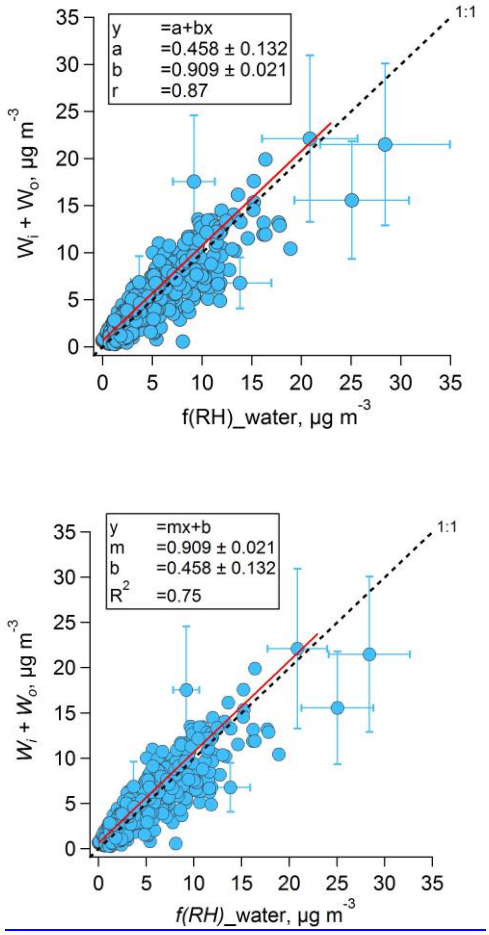


Fig. ~~6~~9. Comparison between total predicted and measured water by nephelometers based on hourly averaged data at CTR (SOAS). An ODR fit was applied. Error bars for selected points are shown.

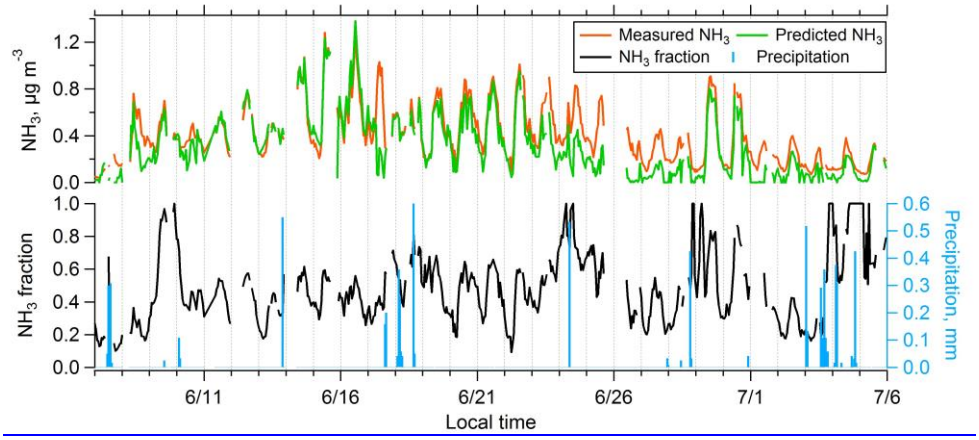
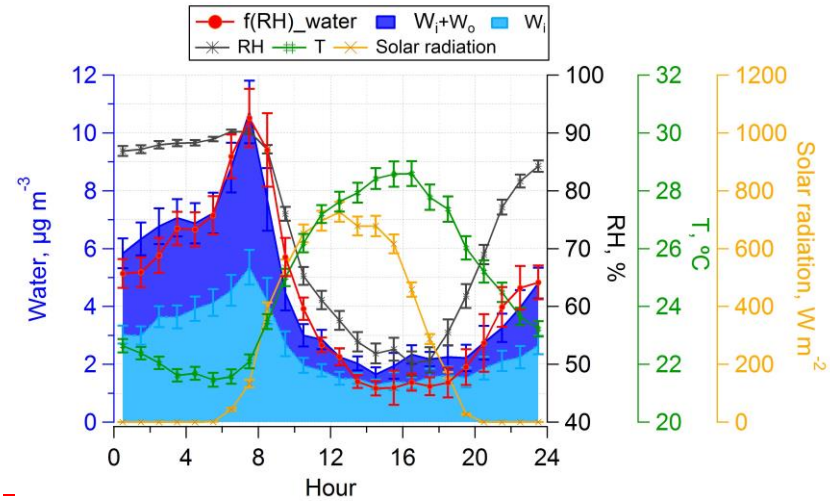
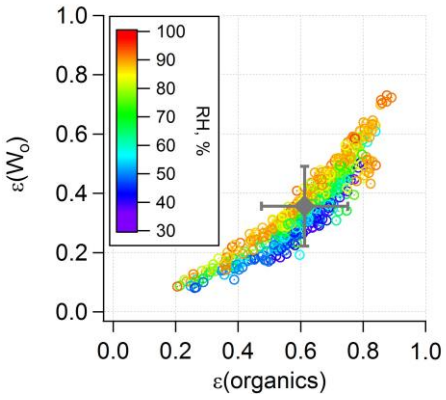

Measured NH3
Predicted NH3
NH3 fraction
Precipitation
NH3, μg m-3
NH3 fraction
Precipitation, mm
6/11
6/16
6/21
6/26
7/1
7/6
Local time

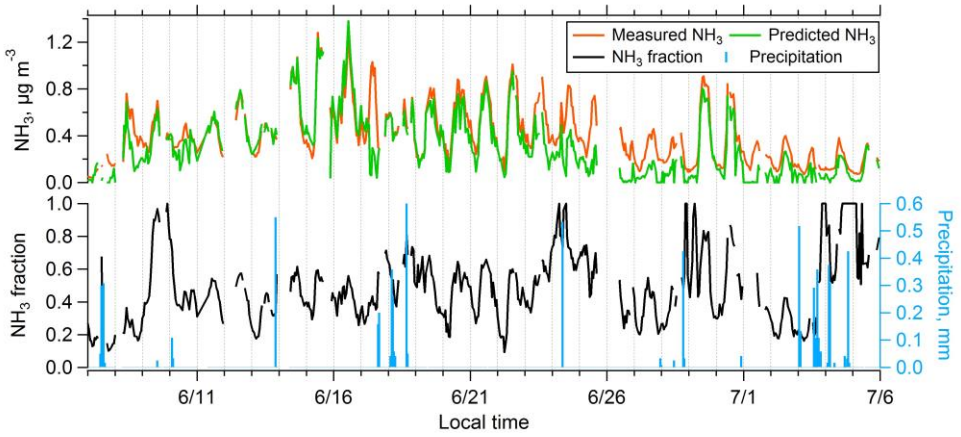
Fig. 10.



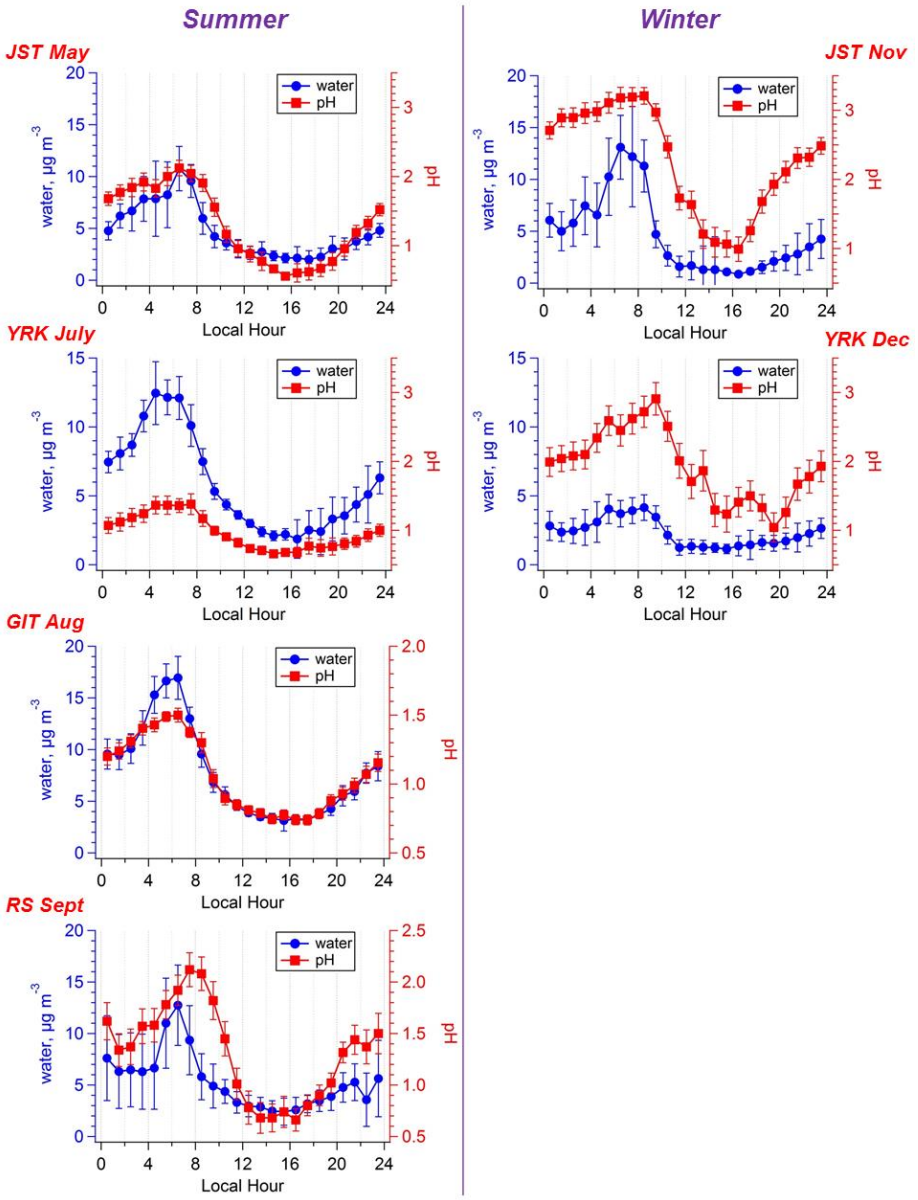
~~**Fig. 7.** CTR (SOAS) diurnal profiles of predicted and measured water, measured RH, T, and solar radiation. Median hourly averages and standard error plotted as error bars at local hour are shown.~~



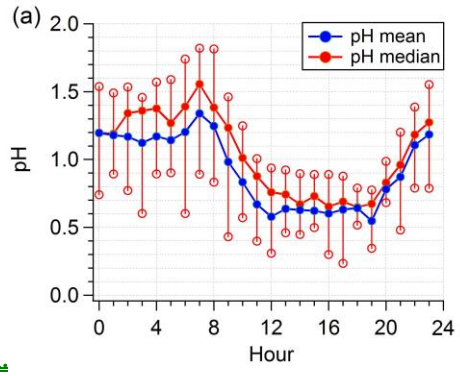
~~Fig. 8.~~ ~~W_o fraction plotted with organic mass fraction at CTR (SOAS).~~ ~~Overall study mean and standard deviation is also shown.~~ ~~$\varepsilon_{organic}$ = 61 ± 14% and ε_{w_o} = 36 ± 14%.~~



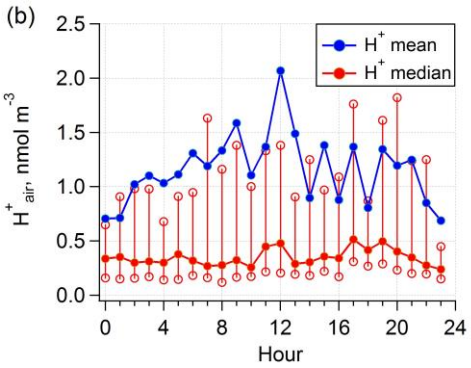
~~Fig. 9.~~ CTR (SOAS) time series of hourly averaged measured $NH_{3(g)}$, predicted $NH_{3(g)}$, $NH_{3(g)}$ fraction (i.e., measured $NH_{3(g)}$/(~~-~~$NH_{3(g)}$+$NH_{4(p)}$~~))~~$^{+}$)) and precipitation.


Summer
Winter
JST May
JST Nov
YRK July
YRK Dec
GIT Aug
RS Sept
water
pH
water, μg m-3
pH
Local Hour

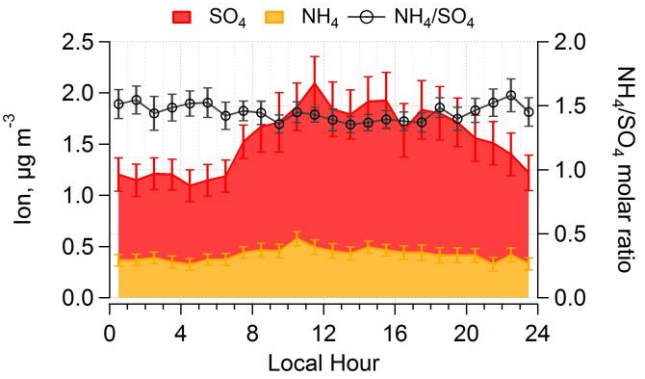
1140



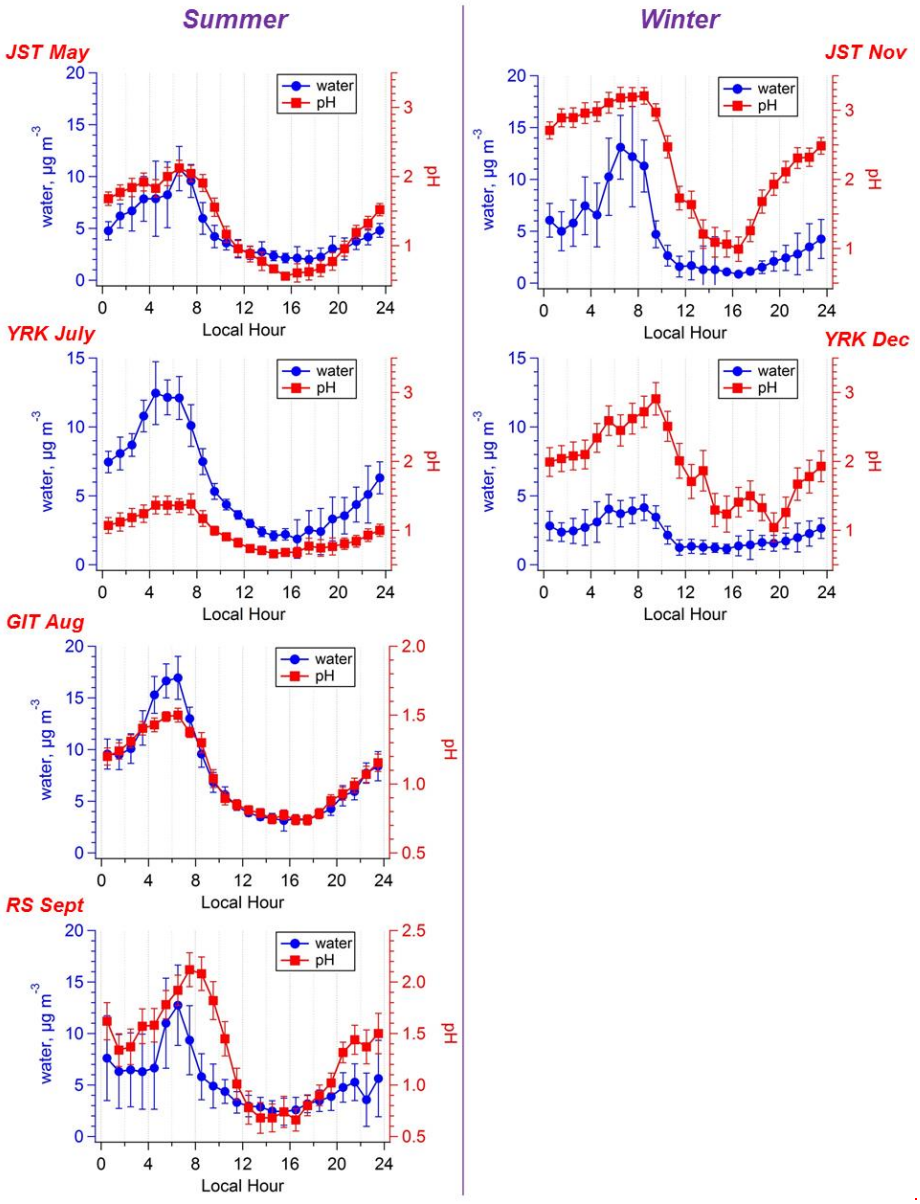
~~Fig. 11.~~



~~Fig. 10.~~ ~~CTR (SOAS) diurnal patterns of calculated pH based on total predicted water ($W_t + W_o$) and H^+_{air} predicted by ISORROPIA-II. Mean and median values are shown, together with 25% and 75 quantiles marked as non-filled circles.~~



~~**Fig. 11.** CTR (SOAS) diurnal patterns of PILS-IC NH_4 and SO_4, and the molar ratio. Median hourly averages and standard error bars at local hours are plotted.~~



~~Fig. 12.~~ LWC and pH diurnal variation at SCAPE sites: comparison between summer and winter. Median hourly averages and standard error bars at local hour are plotted. A bias correction of 1 pH unit is applied due to not considering ammonia partitioning.

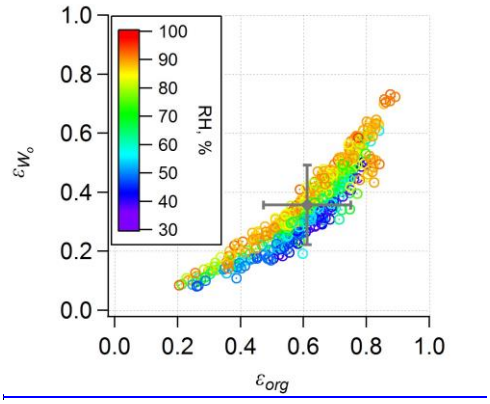
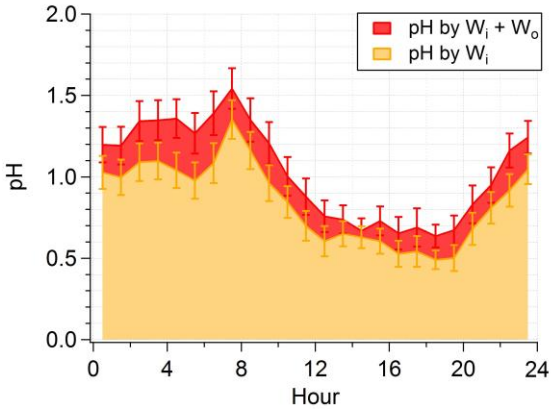

1.0
0.8
0.6
0.4
0.2
0.0
0.0
0.2
0.4
0.6
0.8
1.0
100
90
80
70
60
50
40
30
RH, %
ε_{W_o}
ε_{org}

Fig. 12.

~~Particle water and~~Fine particle pH in the southeastern United States



W_o mass fraction (ε_{W_o}) plotted versus organic mass fraction at CTR (SOAS). Overall study mean and standard deviation is also shown. ε_{Org} = 61 ± 14% and ε_{W_o} = 36 ± 14%.

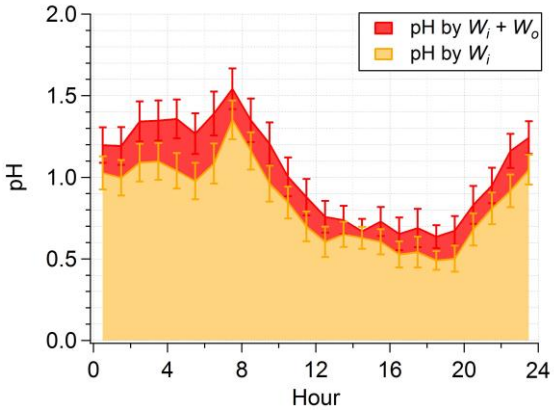
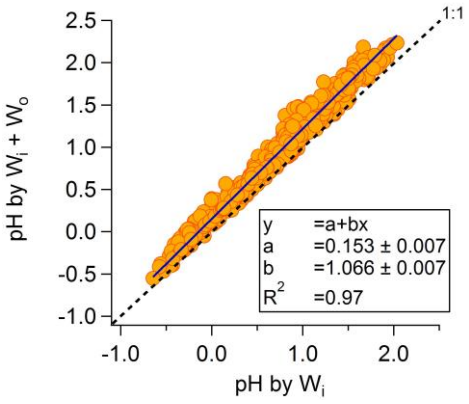


Fig. 13. CTR (SOAS) pH diurnal profiles based on total predicted water and W_i, respectively. Median hourly averages and standard error bars at local hour are plotted.



~~**Fig. 14.** Comparison between pH based on W_t and pH based on total predicted water at CTR (SOAS). An ODR fit was applied.~~