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Letter to the Editor of the ACP Manuscript "An estimation of the $^{18}O/^{16}O$ ratio of UT/LMS ozone based on artefact CO in air sampled during CARIBIC flights" *by* S. Gromov and C. A. M. Brenninkmeijer

S. Gromov (on behalf of all authors)

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Dear Dr. Kaiser,

We are very grateful for your great attention to this work and constructive comments that helped us to improve the quality of this manuscript significantly. Following your suggestions, we have prepared the revised version (please, find the pages with mark-up at the end of this letter). We have addressed all your comments (shown below italicised), on a few of them we have comments or different opinion, as we discuss in the following.

We appreciate very much the time you spent for editing this paper.

Best regards,

Sergey Gromov

Major points:

1) Contamination kinetic framework: <...>

We are not going to carry on further polemics with the Editor who perennially tries to interpret a conceptual expression

$A \rightarrow B \rightarrow C$

as a strict chemical equation (besides that we have never stated that Eq. (A1) is one). We are grateful to the Editor for spending his precious time on enlightening us on textbook kinetics behind which, however, the discussion based on misinterpreted Eq. (A1) has nevertheless no sense. On the other hand, we admit that a mix-up of terms and formulations occurred, which inevitably impedes the Reader in understanding of what we have done. Ultimately, we find that we have rather used improper terminology than the kinetic apparatus, for which we sincerely apologise. We therefore decided to recast the "kinetic framework" (also renaming it to "contamination assessment" for clarity and leaving out any conceptual expressions, *i.e.* Eqs. (A1) and (A2)), as described below. Furthermore, restating the grounds of the kinetic apparatus we use, we are lucky to return to IUPAC definitions and conclude that Eqs. (A2) and (A3) in essence formulate the macroscopic rate of reaction (UIPAC "Gold book", http://dx.doi.org/10.1351/goldbook.O04322), with K and κ being the partial orders of reaction with respect to concentrations of X and O_3, respectively, and k_c in our regression analysis being the so-called "observed" rate coefficient which determines the chemical flux of artefact CO. In the revised version of Appendix A we also better emphasise the set of assumptions involved, namely: (i) reaction time varies negligibly, (ii) changes to abundances of other reactants (if there are any) are negligible, and (iii) changes to other important parameters like temperature, pressure, flight time, *etc.* have no discernible effect on the amount of artefact CO produced. Finally, we state that merely a regression analysis was done.

The "kinetic" framework is further compromised by the injudicious mixing of mole fractions and number densities. You state that the units match on both sides of your equations, but they don't in case of equations (A2) and (A3). E.g., in case of A3, the term [O3]^kappa has units of (cm–3)^(2.06) = cm^(-6.18).

In case of Eq. (A2), it is stated (ll. 367–368): "... (abundances in number density units are used)". Eq. (A3) inherits its units from Eq. (A2) being a reduced from of it. Furthermore, we state (ll. 380–382): "... the product ($\lambda_{O_3} k_c \tau_c$) that proportionates the CO contamination strength and $[O_3]$ is found to be $(5.19 \pm 0.12) \cdot 10^{-5}$ mol/nmol ($\pm 1\sigma$, adj. $R^2 = 0.83$, red. $\chi^2 = 4.0$; mole fraction units are used here for convenience)."

We explicitly state that, for convenience of the Reader, $[O_3]$ (mole fraction of ozone) is to be used with this factor to quantify the CO contamination strength (also in mole fraction units). That is, by multiplying $[O_3]^2$ (nmol/mol)^2 by $(5.19\pm0.12)\cdot10^{-5}$ (mol/nmol) the resulting value is to be in (nmol/mol)^2 × (mol/nmol) = (nmol/mol), *i.e.* in mole fraction units. Such conversion is possible because the pressure and temperature variations do not result in significant variations of air density at the C1 sampling conditions. We neither find "injudicious mixing of mole fractions and number densities" here nor understand how this can "compromise the "kinetic" framework".

I am unable to see how you derive a dependence of the CO artefact formation rate on the squared ozone number density.

This leaves the empirical regression equation to justify the apparent dependence of the artefact mixing ratio on the (approximate) square of the O3 mixing ratio. However, looking at Fig. 1d, I'd suggest that a linear fit would give almost as good a fit as power fit (to the power of 2.06). In addition, a linear fit could be more directly related to a kinetic mechanism (as suggested above).

We appreciate Editor's ability to infer goodness of fit by contemplating the data with a naked eye. Being not confident of applying this method ourselves, however, we used the conventional apparatus (*viz.* R^2 and reduced χ^2 statistic), as duly communicated in Appendix A. We also remain certain, that it is clear to the Editor that the regression analyses were performed sequentially. That is, firstly the expression in Eq. (A3) is regressed with respect to two parameters, namely the product $(\lambda_{O3}\cdot k_c\tau_c)$ and κ, which yields the least value of reduced χ^2 at κ of (2.06±0.38). Secondly, taking (assuming) $\kappa = 2$ the regression of the product $(\lambda_{O3}\cdot k_c\tau_c)$ is done. In this step we also ascertain that regression yields the least reduced χ^2 value at $\kappa = 2$, in contrast to κ being equal any other integer number.

To summarise, we propose the following amendment of Appendix A (and looking forward to your kind suggestions for improvement):

Appendix A. Contamination assessment

We quantify the C1 CO contamination strength (denoted $[CO]_c$, obtained by discriminating the C1 outliers from respective C2 data) in a sequence of regression analyses. We foremost ascertain that no other species or operational parameter (*e.g.* temperature, pressure, flight duration, season, latitude, time of day, *etc.*) measured in C1 appear to determine (*e.g.*, systematically correlate with) $[CO]_c$, except that for $[O_3]$. We hypothesise therefore that a production of artefact CO molecules was initiated by O_3 (via either its decomposition or a reaction with an unknown educt) and proceeded with incorporation of carbon (donated by some carbonaceous species X) and oxygen (donated by O_3 or its derivatives) atoms into final CO. Despite that neither the actual reaction chain nor its intermediates are known, it is possible to describe the artefact CO component produced (hereinafter curly brackets {} denote number densities) as

$$\{CO\}_c = \lambda_{O_3}\, v\, \tau_c\,, \quad \text{(A1)}$$

where the yield λ_{O_3}, a diagnostic quantity, relates the amount of artefact CO molecules produced to the total number of O_3 molecules consumed in the system, τ_c denotes the reaction time (period throughout which sampled air is exposed to contamination), and v stands for the overall rate of the reaction chain. The latter, being regarded macroscopically (empirically), is parameterised to account for the order of reaction chain rate with respect to hypothesised reactants (McNaught and Wilkinson, 1997) as

$$v = k\,\{X\}^K\{O_3\}^\kappa\,, \quad \text{(A2)}$$

where κ and K are the partial orders with respect to X and O_3 number densities, respectively, and k is the rate coefficient. Here it is implied that changes to {X} and $\{O_3\}$ are negligible throughout the exposure time τ_c (typically < 0.1 s for C1 sample line). As stated above, we find that variations in $\{CO\}_c$ correlate exclusively with variations in $\{O_3\}$, hence Eq. (A2) can be reduced by assuming constancy of {X} and K to:

$$v_c = k_c\,\{O_3\}^\kappa\,. \quad \text{(A3)}$$

Here, $k_c = k\{X\}^K$ (often referred to as pseudo-first-order or "observed" rate coefficient) quantifies the rate of reaction chain exclusively propelled by O_3. Finally, using Eqs. (A1) and (A3), the artefact $\{CO\}_c$ component is expressed as

$$\{CO\}_c = b\cdot\{O_3\}^\kappa\,,\ b = \lambda_{O_3} k_c \tau_c \quad \text{(A4)}$$

where the constant proportionality factor b integrates the influence of the unknown (and as we explicate below, likely invariable) {X}, k, K and τ_c.

Eq. (A4) defines the regression expression using which we attempt to fit the values of $\{CO\}_c$ as a function of κ, $\{O_3\}$ and b. In the first regression iteration we keep both κ and b as free parameters, which provides best approximation at $\kappa = 2.06 \pm 0.38$, suggesting reactions of two O_3 molecules in case elementary reactions constitute the reaction mechanism, or two elementary steps involving O_3 or its derivatives in case a stepwise reaction is involved (McNaught and Wilkinson, 1997). In a subsequent regression iteration we set $\kappa = 2$, which yields better (as opposed to the first iteration) estimate of b of $(5.19 \pm 0.12) \cdot 10^{-5}$ mol/nmol ($\pm 1\sigma$, adj. $R^2 = 0.83$, red. $\chi^2 = 4.0$; here the equivalent value in mole fraction units is quoted for the convenience of relating fitted $[CO]_c$ and $[O_3]^2$). At last, we ascertain that the best regression results are obtained particularly at $\kappa = 2$, as indicated by the regression statistic (R^2 and χ^2) that asymptotically improves when a set of regressions with neighbouring (*i.e.* below and above 2) integer values of κ is compared. The low uncertainty (within ±3%) associated with the estimate of b confirms an exclusive dependence of the contamination source on the O_3 mixing ratio, as well as much similar reaction times τ_c. The regressed value of $[CO]_c$ as a function of $[O_3]$ is presented in Fig. 1 (d) (solid line). It is possible to constrain the overall yield λ_{O_3} of CO molecules in the artefact source chain to be between 0.5 and 1, comparing the magnitude of $[CO]_c$ to the discrepancy between the $[O_3]$ measured in C1 and C2 (±20 nmol/mol, taken equal to the $[O_3]$ bin size owing to the $N_2O{-}O_3$ and $H_2O{-}O_3$ distributions matching well between the datasets). Lower λ_{O_3} values, otherwise, should have resulted in a noticeable (*i.e.*, greater than 20 nmol/mol) decrease in the C1 O_3 mixing ratios with respect to the C2 levels.

2) Please do not deviate unnecessarily from established nomenclature, e.g. "abundance" is not a defined concentration quantity and should not be used as such. It is only defined in terms of isotopic abundances. Please stick with the terms "mixing ratio" (and explain at some point that you understand this to mean "mole fraction") and, if required, "number density". You might find the following paper of interest:

Cvitaš T (1996) Quantities describing compositions of mixtures Metrologia:35 doi:10.1088/0026-1394/33/1/5

To avoid confusion, please use the same symbol for a specific quantity throughout the manuscript. For example, you use both square brackets and "C" for mixing ratios. Please use only one of these two. Also, if you use square brackets to denote mixing ratios, please do not also use this for number densities. Note that the IUPAC Green Book recommends the symbol "y" for mole fractions in gaseous mixtures and the symbol "C" for number densities (http://media.iupac.org/publications/books/gbook/IUPAC-GB3-2ndPrinting-Online-22apr2011.pdf).

We replace all “C” with equivalent “[CO]” terms throughout the revised manuscript. For the number densities (used only in Appendix A) we prefer to use curly brackets, *e.g.* “{CO}”.

3) Please decide on your preferred delta notation and stick with it. If you were to follow IUPAC recommendations, you would have to write delta(18O, O3), delta(18O, CO), etc. However, I acknowledge that, unfortunately, many isotope geochemists have adopted the incorrect notation delta18O(O3) etc., perhaps because they mixed up the physical quantity symbol delta with a mathematical operator (which, however, would also be wrong because a mathematical operator cannot be applied to a chemical symbol).

We prefer to stick with the "$\delta^{18}O(O_3)$" notation, as we acknowledge that the vast majority of literature using delta notation sticks with it, in contrast to the UIPAC recommendations.

4) Please remove tilde signs and round values adequately and/or give uncertainties, as appropriate. In any case, the tilde sign should not be used to indicate approximations; the correct symbol is "≈" (two wavy lines one above the other). However, since very few of the values mentioned in the paper are exact, the approximation sign is almost always redundant and would be best omitted. Instead the corresponding value should be given with an appropriate number of significant figures according to its uncertainty.

We have corrected all occurrences according to your kind suggestions.

Technical corrections:

5: A photochemical nature is not ascertained because no detailed process studies or kinetic modelling has been undertaken. Photochemistry is unlikely in the (largely) dark pipes of the sampling system. Rather, you "suggest that the magnitude of the artefact is proportional to the square of the ozone mixing ratio".

Correct. We change it to “(ii) ascertain the chemical nature and quantify the strength of the contamination, and ...”.

8: Replace "signature" with "delta" and 18O/16O ratios with delta18O. Replace "unambiguously" with "likely".

We use “isotope composition” instead of “isotope signature”.

14: Delta needs to be in italics.

We are certain this will be taken care of during the typesetting of the manuscript (if it comes to that) according to the ACP conventions (these imply italicised delta, as far as we know).

93: Add "in situ" after C2 for clarity.

Added. In Sect. 2.1 (above) we already state that, however.

99: "in NH tropospheric emissions" (otherwise this would be a tautology)

We see no tautology here. The CO variations result from mixing of the little varying stratospheric [CO] and largely varying tropospheric [CO]. It is the result of mixing we discuss here. Besides, variations in tropospheric [CO] are *by far* more strongly determined by the presence of hydroxyl radical that by the variations in emissions.

102: "in C1 and C2 [CO], for [O3] > 400 nmol/mol the C1 CO mixing ratios [...]"

This comment is unclear to us. We describe continuous changes in [CO] with increasing [O_3], this will change the meaning of the sentence to something we do not intend to state.

104: "In the 580-600 nmol/mol [O3] bin"

This comment is unclear to us. We describe to what [CO] in C1 one observes in particular bin (around 580 nmol/mol of [O_3]), this will change the meaning of the sentence to something we do not intend to state.

105: "accommodates and extra 14 nmol/mol"

Here we meant that this [CO] contains extra 15 nmol/mol as compared to average C2 value. We adjust the statement accordingly.

160-162: This sentence duplicates the message of the previous one and can be deleted.

Please explain. The statements “[CO] from WAS and *in situ* measurements correlate well” and “anomalies in both [CO] and $\delta^{18}O(CO)$ manifest functions of [O_3]” do not appear duplicate to us.

171: This ratio appears to be wrong and should be 600:70 based on the values given here.

Rather the ratio of 600:70 appears to be wrong.

Taking the weighted sum of 700 (85%) and 60 (15%), one receives 700*0.85+60*0.15 = 604. Taking the weighted sum of 24 (85%) and 125 (15%), one receives 24*0.85+125*0.15 = 39.15. This is 4% and 2% off from the quoted 580 and 40, respectively. The ratio 600:70 implies the resulting [CO] about 80% higher than the correct value.

To be more precise, we adjust the figures accordingly:

> For example, mixing of two air parcels in a 16%:84% proportion (by moles of air) with typical [O_3]:[CO] of 700:24 (stratospheric) and 60:125 (tropospheric), respectively, yields an integrated composition with [O_3]:[CO] of 598:40 which indeed corresponds to C1 data (this case is exemplified by the mixing curve in Fig. 1).

184: This sentence does not make sense. [H2O] cannot witness [CO]:[O3] ratios. Please change or delete the entire sentence from l. 183 to 185 as it does not appear to add any information.

Thank you, we change “[CO]:[O_3]” to “[O_3]:[H_2O]” for clarity.

197-200: This sentence is unclear - is this meant to be a contradiction, confirmation, fact, assumption, hypothetical calculation or what?

We change it to:

> Our simulations of the ‘translational mixing’ effects confirm that the actual C2 CO−O_3 distribution in the region of interest ([O_3] of 540−620 nmol/mol) remains insensitive to averaging intervals of up to 300 s.

206: Replace "abundance" with "mixing ratio" and change "imply" to "suggest". Why photochemical? O3 could reaction in the dark as well?

We change it to "O_3-mediated production of CO took place".

223: Please delete "Practically" and change "resort" to "use". The Keeling plot itself does _not_ require an estimate of [CO]_c; however, your data selection criterion (for delta_true) does. Please change this sentence accordingly.

Perhaps, the Editor has misunderstood the message of the sentence. Here we emphasise that we can employ the MM using solely the estimate of the contamination strength (*i.e.*, the amount of molecules admixed to the reservoir with some initial composition). Furthermore, do you imply that using the Keeling plot one does not require to know the amount of molecules admixed into a reservoir with known starting composition? (It obviously would be nonsense, of course, perhaps we did not understand your comment?)

227: Please break this into two equations and number the equations. Indices should not be italics. The symbol C should be replaced with [CO] for consistency (or "y" if you adopt IUPAC symbols; see comment 2 above).

See our answer to comment 2) above. Regarding the italics, see our answer to l. 14 above.

253: The symbol 13delta_c has not been defined. For consistency, this should be delta13C_c(CO), or, following conventional symbol and index notation, delta_c(13C, CO).

The Editor contradicts himself here. In the previous version of the manuscript we used a consistent notation using indices to distinguish δ_c for ^{13}C and ^{18}O, which the Editor requested to remove (see the comment on l. 227 of the previous version). Since distinguishing different δ_c, δ_a and δ_t variables is obviously necessary we return to the previous notation, *e.g.* $^{13}\delta_c$ and $^{18}\delta_c$.

348: Add "in combination with an empirical parameterisation of the [CO] artefact in terms of the O3 mixing ratio" after brackets, followed by "to single out ..."

We would like to keep the current formulation, as we already make a statement above (ll. 345−346) on the quantification of the artefact CO production.

Fig. 6: The x-axis label should be "MM", not MMA. The legend labels should be delta18O_c(O3) and delta13C_c(O3); also in the caption.

We change the labels to $^{18}\delta_c$ and $^{13}\delta_c$, respectively, that are clearly associated with calculations with the MM. This also allows to avoid somewhat confusing $\delta^{13}C_c(O_3)$ (the carbon isotope ratio from O_3 makes no sense here).

613: "18Odelta_t" should be deleted.

Perhaps, the Editor did not understand this statement. Different R^2 values are obtained for different signatures being regressed, *i.e.* for $^{18}\delta_c$ and $^{13}\delta_c$. Here we emphasise the pair of $^{13}\delta_c$ values encircled corresponds to the pair of best-guessed $^{18}\delta_c$ values which are obtained with highest R^2 value.

An estimation of the $^{18}O/^{16}O$ ratio of UT/LMS ozone based on artefact CO in air sampled during CARIBIC flights

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Abstract

An issue of O_3-driven artefact production of CO in the upper troposphere/lowermost stratosphere (UT/LMS) air analysed in the CARIBIC−1 project is being discussed. By confronting the CO mixing and isotope ratios obtained from different analytical instrumentation, we (i) reject natural/artificial sampling and mixing effects as possible culprits of the problem, (ii) ascertain the chemical nature and quantify the strength of the contamination, and (iii) demonstrate successful application of the isotope mass-balance calculations for inferring the isotope composition of the contamination source. The $\delta^{18}O$ values of the latter indicate the oxygen likely being inherited from O_3. The $\delta^{13}C$ values hint at reactions of trace amounts of organics with stratospheric O_3 that could have yielded the artificial CO. While the exact contamination mechanism is not known, it is clear that the issue pertains only to the earlier (first) phase of the CARIBIC project. Finally, estimated UT/LMS ozone $\delta^{18}O$ values are lower than those observed in the stratosphere within the same temperature range, suggesting that higher pressures (240−270 hPa) imply lower isotope fractionation controlling the local $\delta^{18}O(O_3)$ value.

1 Introduction

[1] Accurate determination of the atmospheric carbon monoxide (CO) mixing ratio based on the collection of air samples depends on the preservation of the mixing ratio of CO inside the receptacle, from the point of sampling to the moment of physicochemical analysis in a laboratory. A well known example in our field of research is the filling of pairs of glass flasks at South Pole

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Station for analysis at NOAA in Boulder, Colorado, USA (Novelli *et al.*, 1998). There, the duplicate air sampling allowed for a degree of quality control which in view of the long transit times, especially during polar winter, was a perhaps not perfect, but certainly a practical measure. Here we deal with a different case: Using aircraft-based collection of very large air samples rendered duplicate sampling unpractical, yet analyses could be performed soon after the sampling had taken place because of the proximity of the aircraft's landing location to the laboratory involved. A presumption of the analytical integrity of the process was that the growth of CO in receptacles is gradual and takes its time. Reminding Thomas Henry Huxley's statement, "The great tragedy of Science – the slaying of a beautiful hypothesis by an ugly fact", it turned out, however, that for air we collected in stainless steel tanks in the upper troposphere/lowermost stratosphere (UT/LMS) higher CO values were measured in the laboratory than measured *in situ* during the collection of these air samples. Moreover, measurement of the stable oxygen isotopic composition of CO from these tanks revealed additional isotopic enrichments in ^{18}O of 10‰ or more. It was soon realised that this phenomenon was due to the formation of CO in these tanks and/or possibly in the sampling system and inlet tubing used, by reactions involving ozone (Brenninkmeijer *et al.*, 1999).

[2] Unexpectedly high $^{18}O/^{16}O$ ratios in stratospheric ozone (O_3) were discovered by Konrad Mauersberger using a balloon-borne mass spectrometer (Mauersberger, 1981), which has triggered a series of theoretical and experimental studies on atmospheric O_3 heavy isotope enrichments (see, *e.g.*, Schinke *et al.* (2006) for a review). In view of the advances in theoretical and laboratory studies on the isotopic composition of O_3 atmospheric measurements are welcome, they do however form a challenge. In the stratosphere O_3 number concentrations are high, but the remoteness of the sampling domain is a problem. In the troposphere, low O_3 number densities are the main obstacle, as indicated by few experiments performed to date (Krankowsky *et al.*, 1995; Johnston and Thiemens, 1997; Vicars and Savarino, 2014). Nevertheless, recent analytical improvements, namely the use of an indirect method of reacting atmospheric O_3 with a substrate that can be analysed for the isotopic composition of the O_3-derived oxygen (Vicars *et al.*, 2012), has greatly improved our ability to obtain information on the O_3 isotopic composition.

[3] Although the increase of CO concentrations in air stored in vessels is a well recognised problem, to our knowledge a specific O_3-related process has not been reported yet. Here we discuss this phenomenon and turn its disadvantage into an advantage, namely that of obtaining an estimate of the oxygen isotopic composition of O_3 in the UT/LMS, an atmospheric domain not

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yet covered by specific measurements. The air samples we examine in this study were collected onboard a passenger aircraft carrying an airfreight container with analytical and air/aerosol sampling equipment on long distance flights from Germany to South India and the Caribbean within the framework of the CARIBIC project (Civil Aircraft for the Regular Investigation of the atmosphere Based on an Instrument Container, http://www.caribic-atmospheric.com).

2 Experimental and results

2.1 Whole air sampling

[4] CARIBIC−1 (Phase #1, abbreviated hereafter "C1") was operational from November 1998 until April 2002 using a *Boeing* 767-300 ER operated by LTU International Airlines (Brenninkmeijer *et al.*, 1999). Using a whole air sample (WAS) collection system, twelve air samples were collected per flight (of 8−10 hours duration at cruise altitudes of 10−12 km) in stainless steel tanks for subsequent laboratory analysis of the mixing ratios (*i.e.* mole fractions) of various trace gases, including ^{14}CO. Large air samples were required in view of the ultra-low number density of this mainly cosmogenic tracer (10−100 molecules cm^{-3} standard temperature and pressure (STP), about 0.4−4 amol/mol). Hereinafter STP denotes dry air at 273.15 K, 101325 Pa. Each C1 WAS sample (holding 350 litres of air STP) was collected over 15−20 min intervals representing the number density-weighted average of the compositions encountered along flight segments of about 250 km. The overall uncertainty of the measured WAS CO is less than ±1% for the mixing ratio and ±0.1‰/±0.2‰ for $\delta^{13}C(CO)/\delta^{18}O(CO)$, respectively (Brenninkmeijer, 1993; Brenninkmeijer *et al.*, 2001). Isotope compositions are reported throughout this manuscript using the so-called delta value $\delta = (R/R_{st}-1)$ relating the ratio R of rare (^{13}C, ^{18}O or ^{17}O) over abundant isotopes of interest to the standard ratio R_{st}. These are Vienna Standard Mean Ocean Water for $^{18}O/^{16}O$ (Gonfiantini, 1978; Coplen, 1994) and $^{17}O/^{16}O$ (Assonov and Brenninkmeijer, 2003), and Vienna Pee Dee Belemnite for $^{13}C/^{12}C$ (Craig, 1957), respectively. As we mention above, the oxygen isotope composition of the CO present in these WAS samples was corrupted, in particular when O_3 levels were as high as 100−600 nmol/mol.

[5] CARIBIC−2 (Phase #2, referred to as "C2") started operation in December 2004 with a Lufthansa *Airbus* A340-600 fitted with a new inlet system and air sampling lines, including perfluoroalkoxy alkane (PFA) lined tubing for trace gas intake (Brenninkmeijer *et al.*, 2007). No flask CO mixing/isotope ratio measurements are performed in C2.

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2.2 On-line instrumentation

[6] In addition to the WAS collection systems, both C1 and C2 measurement setups include different instrumentation for on-line detection of [CO] and [O_3] (hereinafter the squared brackets [] denote the mixing ratio of the respective species). *In situ* CO analysis in C1 is done using a gas chromatography (GC)-reducing gas analyser which provides measurements every 130 s with an uncertainty of ±3 nmol/mol (Zahn *et al.*, 2000). In C2, a vacuum ultraviolet fluorescence (VUV) instrument with lower measurement uncertainty and higher temporal resolution of ±2 nmol/mol in 2 s (Scharffe *et al.*, 2012) is employed. Furthermore, the detection frequency for O_3 mixing ratios has also increased, *viz.*, from 0.06 Hz in C1 to 5 Hz in C2 (Zahn *et al.*, 2002; Zahn *et al.*, 2012).

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2.3 Results

[7] When comparing the CO mixing ratios in relation to those of O_3 for C1 and C2, differences are apparent in the LMS, where C2 [CO] values are systematically lower. This is illustrated in Fig. 1 (a) which presents the LMS CO-O_3 distribution of the C2 *in situ* measurements overlaid with the C1 *in situ* and WAS data. The entire C1 CO/O_3 dataset is presented in Fig. 2. For the *in situ* CO datasets we calculated the statistics (Fig. 1 (b)) of the samples with respective O_3 mixing ratios clustered in 20 nmol/mol bins, *i.e.* the median and spread of [CO] as a function of [O_3] analysed. The interquartile range, IQR, is used in the current analysis as a robust measure of the data spread instead of the standard deviation. The data exhibit large [CO] variations at [O_3] below 400 nmol/mol that primarily reflect pronounced seasonal variations in the NH tropospheric CO mixing ratio. With increasing [O_3], [CO] decreases to typical stratospheric values, and its spread reduces to mere 3.5 nmol/mol and less, as [O_3] surpasses 500 nmol/mol. Despite the comparable spread in C1 and C2 [CO], from 400 nmol/mol of [O_3] onwards the C1 CO mixing ratios start to level off, with no samples below 35 nmol/mol having been detected, whereas the C2 levels continuously decline. By the 580 nmol/mol O_3 bin, C1 [CO] of $39.7^{+0.7}_{-0.3}$ nmol/mol contains some extra 15 nmol/mol compared to $25.6^{+1.2}_{-1.1}$ nmol/mol typical for C2 values. Overall, at [O_3] above 400 nmol/mol the conspicuously high [CO] is marked in about 200 *in situ* C1 samples, of which 158 and 69 emerge as statistically significant mild and extreme outliers, respectively, when compared against the number of C2 samples ($n > 3 \cdot 10^5$). The conventions here follow Natrella (2003), *i.e.* ±1.5 and ±3 IQR ranges define the inner and outer statistical fences (ranges outside which the data points are considered mild and extreme outliers) of the C2 [CO] distribution in every O_3 bin, respectively. The statistics include the samples in bins with average

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[O_3] of 420−620 nmol/mol. None of C1 CO at [O_3] above 560 nmol/mol agrees with the C2 observations. Because the CO-O_3 distribution cannot have changed over the period in question, we find that an apparent relative excess CO of up to 55% justifies and investigation into sampling artefacts and calibration issues.

[8] Unnatural elevations in $\delta^{18}O(CO)$ from WAS measurements are also evident, as shown in Figs. 3 and 4. The large $\delta^{18}O(CO)$ elevations that reach beyond +16‰ are found to be proportional to the concomitant O_3 mixing ratios (denoted with colour) and are more prominent at lower [CO]. Lower $\delta^{18}O(CO)$ values, however, are expected based on our knowledge of UT/LMS CO sources (plus their isotope signatures) and available *in situ* observations (Fig. 3, shown with triangles), as elucidated by Brenninkmeijer *et al.* (1996) (hereafter denoted as "B96"). That is, the greater the proportion of stratospheric CO, the greater its fraction stemming from methane oxidation with a characteristic $\delta^{18}O$ of 0‰ or lower (Brenninkmeijer and Röckmann, 1997). This occurs because the CO sink at ruling UT/LMS temperatures proceeds more readily than its production, as the reaction of hydroxyl radical (OH) with CO, being primarily pressure-dependent, is faster than the temperature-sensitive reaction of OH with CH_4. Furthermore, as the lifetime of CO quickly decreases with altitude, transport-mixing effects take the lead in determining the vertical distributions of [CO] and $\delta^{18}O(CO)$ above the tropopause, hence their mutual relationship. This is seen from the B96 data at [CO] below 50 nmol/mol that line-up in a near linear relationship towards the end-members with lowest $^{18}O/^{16}O$ ratios. These result from the largest share of the ^{18}O-depleted photochemical component and extra depletion caused by the preferential removal of $C^{18}O$ in reaction with OH (fractionation about +11‰ at pressures below 300 hPa, Stevens *et al.*, 1980; Röckmann *et al.*, 1998b).

[9] We are confident that the enhancements of C1 $C^{18}O$ originate from O_3, whose large enrichment in ^{18}O (above +60‰ in $\delta^{18}O$, Brenninkmeijer *et al.*, 2003) is typical and found transferred to other atmospheric compounds (see Savarino and Morin (2012) for a review). In Fig. 3 it is also notable that not only the LMS compositions are affected but elevations of (3−10)‰ from the bulk $\delta^{18}O(CO)$ values are present in more tropospheric samples with [CO] of up to 100 nmol/mol. These result from the dilution of the least affected CO-rich tropospheric air by CO-poor, however substantially contaminated, stratospheric air, sampled into the same WAS tank. Such sampling-induced mixing renders an unambiguous determination of the artefact source' isotope signature rather difficult, because neither mixing nor isotope ratios of the admixed air portions are known sufficiently well (see below).

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[10] Differences between the WAS and *in situ* measured [CO] – a possible indication that the $\delta^{18}O(CO)$ contamination pertains specifically to the WAS data – average at $\overline{\Delta}$(WAS−*in situ*) = (5.3±0.2) nmol/mol (±1 standard deviation of the mean, n = 408) and happen to be random with respect to any operational parameter or measured characteristic in C1, *i.e.* irrespective of CO or O_3 abundances. The above mentioned discrepancy remained after several calibrations between the two systems had been performed, and likely results from the differences in the detection methods, drifts of the calibration standards used (see details in Brenninkmeijer *et al.*, 2001) and a short-term production of CO in the stainless steel tanks during sampling. The large spread of Δ(WAS−*in situ*) of ±3.5 nmol/mol (±1σ of the population) ensues from the fact that the *in situ* sampled air corresponds to (2−4)% of the concomitantly sampled WAS volume, as typically 6−7 *in situ* collections of 5 s were made throughout one tank collection of 17−21 min. The integrity of the WAS CO is further affirmed by the unsystematic distribution of the artefact compositions among tanks (in contrast to that for $\delta^{18}O(CO_2)$ in C1 discussed by Assonov *et al.*, 2009). Overall, the WAS and *in situ* measured CO mixing ratios correlate extremely well (adj. R^2 = 0.972, slope of 0.992±0.008 (±1σ), n = 408). However, both anomalies in [CO] and $\delta^{18}O(CO)$ manifest clear but complex influences of the concomitant [O_3]. That is, the C1 *in situ* and WAS data very likely evidence artefacts pertaining to the O_3-driven effect of the same nature. Below we discuss and quantify these influences.

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3 Discussion

[11] Three factors may lead to the (artefact) distributions seen for C1 *in situ* [CO] at LMS O_3 mixing ratios, namely:

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[12] (i) Strong (linear) natural mixing, such as enhanced stratosphere-troposphere exchange (STE), when a [CO] outside the statistically expected range results from the integration of air having dissimilar ratios of the tracers' mixing ratios, *viz.* [O_3]:[CO]. For example, mixing of two air parcels in a 16%:84% proportion (by moles of air) with typical [O_3]:[CO] of 700:24 (stratospheric) and 60:125 (tropospheric), respectively, yields an integrated composition with [O_3]:[CO] of 598:40 which indeed corresponds to C1 data (this case is exemplified by the mixing curve in Fig. 1). Nonetheless, occurrences of rather high stratospheric CO mixing ratios (in our case, 40 nmol/mol at the concomitant [O_3] of 500−600 nmol/mol compared to the typical 24−26 nmol/mol) are rare. For instance, a deep STE similar to that described by Pan *et al.* (2004) was observed by C2 only once (*cf.* the outliers at [O_3] of 500 nmol/mol in Fig. 1), whereas the C1 outliers were exclusively registered in some 12 flights during

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1997–2001. No relation between these outliers and the large-scale [CO] perturbation due to extensive biomass burning in 1997/1998 (Novelli *et al.*, 2003) is established, otherwise elevated CO mixing ratios should manifest themselves at lower [O_3] as well. Other tracers detected in CARIBIC provide supporting evidence against such strongly STE-mixed air having been captured by C1. That is, the binned distributions for water vapour and de-trended N_2O mixing ratios (not shown here) are similar for C1 and C2. Whereas the small relative variations in atmospheric [N_2O] merely confirm matching [O_3] distributions in CARIBIC, the stratospheric [H_2O] distributions witness no [O_3]:[H_2O] values corresponding to those of the C1 outliers, suggesting the latter being unnaturally low.

[13] (ii) Mixing effects can also occur artificially, originating from sampling peculiarities or data processing. Since the CARIBIC platform is not stationary, about 5 s long sampling of an *in situ* air probe in C1 implies integration of the air compositions encountered along some hundred metres, owing to the high aircraft speed. This distance may cover a transect between tropospheric and stratospheric filaments of different compositions. The effect of such 'translational mixing' can be simulated by averaging the sampling data with higher temporal frequency over longer time intervals. In this respect, the substantially more frequent CO data in C2 (sampling interval <1 s) were artificially averaged over a set of increasing intervals to reckon whether the long sampling period in C1 could be the culprit for skewing its CO–O_3 distribution. As a result, the original C2 data and their averages (equivalent to the C1 CO sample injection time) differ negligibly, as do the respective [O_3]:[CO] values. Our simulations of the 'translational mixing' effects confirm that the actual C2 CO–O_3 distribution in the region of interest ([O_3] of 540–620 nmol/mol) remains insensitive to averaging intervals of up to 300 s. Furthermore, a very strong artificial mixing with an averaging interval of at least 1200 s (comparable to C1 WAS sampling time) is required to yield the averages from the C2 data with [O_3]:[CO] characteristic for the C1 outliers.

[14] (iii) In view of the above, it is unlikely that any natural or artificial mixing processes are involved in the stratospheric [CO] discrepancies seen in C1. We therefore conclude that the sample contamination in C1 occurred prior to the probed air reaching the analytical instrumentation and WAS sampling tanks in the container, since clearly elevated stratospheric CO mixing ratios are common to WAS and *in situ* data. Two more indications, *viz.* growing [CO] discrepancy with increasing O_3 abundance, and the strong concomitant signal in $\delta^{18}O(CO)$, suggest that O_3-mediated production of CO took place. Further, by confronting the C1 and C2 [CO] measure-

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ments in a regression analysis (detailed in Appendix A), we quantify the artefact component $[CO]_c$ being chiefly a function of O_3 mixing ratio as

$$[CO]_c = b \cdot [O_3]^2,\ b = (5.19 \pm 0.12) \cdot 10^{-5}\ [\text{mol/nmol}], \quad (1)$$

which is equivalent to 8−18 nmol/mol throughout the respective $[O_3]$ range of 400−620 nmol/mol (see Fig. 1 (d)). Subtracting this artefact signal yields the corrected *in situ* C1 CO−O_3 distribution conforming to that of C2 (*cf.* red symbols in Fig. 1 (a)).

[15] Importantly, since we can quantify the contamination strength using only the O_3 mixing ratio, the continuous *in situ* C1 $[O_3]$ data allow estimating the integral artefact CO component in each WAS sample and, if the isotope ratio of contaminating O_3 is known, to derive the initial $\delta^{18}O(CO)$. The latter, as it was mentioned above, is subject to strong sample-mixing effects, which is witnessed by $\delta^{18}O(CO)$ outliers even at relatively high [CO] up to 100 nmol/mol. Accounting for such cases is, however, problematic since it is necessary to distinguish the proportions of the least modified (tropospheric) and significantly affected (stratospheric) components in the resultant WAS sample mix. Since this information is not available, we applied an *ad hoc* correction approach, as described in the following. This approach is capable of determining the contamination source (*i.e.*, O_3) isotope signature as well.

3.1 Contamination isotope signatures

[16] We use the differential mixing model (MM, originally known as the "Keeling-plot"), because it requires only the estimate of the artefact component mixing ratio, but no assumptions on the (unknown) shares and isotope signatures of the air portions mixed in a given WAS tank. The MM parameterises the admixing of the portion of artefact CO to the WAS sample with the "true" initial composition, as formulated below:

$$[CO]_a = [CO]_t + [CO]_c \quad (2)$$

$$^i\delta_a [CO]_a = {}^i\delta_t [CO]_t + {}^i\delta_c [CO]_c \quad (3)$$

where indices *a*, *c* and *t* distinguish the mixing ratios and isotope compositions $^i\delta$ ($^{18}\delta$ and $^{13}\delta$ for ^{13}C and ^{18}O, respectively) pertaining to the analysed sample, estimated contamination and "true" composition sought (*i.e.*, $[CO]_t$ and $^i\delta_t$), respectively. Here the contamination strength $[CO]_c$ is derived by integrating Eq. (1) using the *in situ* C1 $[O_3]$ data for each WAS sample. By rewriting the above equation with respect to the isotope signature of the analysed CO, one obtains:

$$^i\delta_a = {}^i\delta_c + ({}^i\delta_t - {}^i\delta_c)\,[CO]_c/[CO]_a, \quad (4)$$

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which signifies that linear regression of $^{i}\delta_a$ as a function of the reciprocal of $[CO]_a$ yields the estimated contamination signature $^{i}\delta_c$ at $([CO]_a)^{-1} \rightarrow 0$ when invariable "true" compositions ($[CO]_t$, $^{i}\delta_t$) are taken (the Keeling plot detailing these calculations is shown in Fig. 5). We therefore apply the MM described by Eq. (2) to the subsets of samples picked according to the same reckoned $[CO]_t$ (within a ±2 nmol/mol window, $n > 7$). Such selection, however, may be insufficient: Due to the strong sampling effects in the WAS samples (see previous Section), it is possible to encounter samples that integrate different air masses to the same $[CO]_t$ but rather different average $^{i}\delta_t$. The solution in this case is to refer to the goodness of the MM regression fit, because the R^2 intrinsically measures the linearity of the regressed data, *i.e.* closeness of the "true" values in a regarded subset of samples, irrespective of underlying reasons for that.

[17] Higher R^2 values thus imply higher consistency of the estimate, as demonstrated in Fig. 6 showing the calculated $^{i}\delta_c$ for $[CO]_t$ below 80 nmol/mol as a function of the regression R^2. The latter decreases with greater $[CO]_t$ (*i.e.*, larger sample subset size, since tropospheric air is more often encountered) and, correspondingly, larger variations in $^{i}\delta_t$. Ultimately, at lower R^2 the inferred $^{18}\delta_c$ converge to values slightly above zero expected for uncorrelated data, *i.e.* C1 $\delta^{18}O(CO)$ tropospheric average. A similar relationship is seen for the $^{13}\delta_c$ values (they converge around −28‰), however, there are no consistent estimates found (R^2 is generally below 0.4). Since such is not the case for $\delta^{18}O$, the MM is not sufficiently sensitive to the changes caused by the contamination, which implies that the artefact CO $\delta^{13}C$ should be within the range of the "true" $\delta^{13}C(CO)$ values. Interestingly, the MM is rather responsive to the growing fraction of the CH_4-derived component in CO with increasing $[O_3]$, as the $^{13}\delta_c$ value of −(47.2±5.8)‰ inferred at R^2 above 0.4 is characteristic for the $\delta^{13}C$ of methane in the UT/LMS. It is important to note that we have accounted for the biases in the analysed C1 WAS $\delta^{13}C(CO)$ expected from the mass-independent isotope composition of O_3 (see details in Appendix B).

[18] We derive the "best-guess" estimate of the admixed CO ^{18}O signature at $^{18}\delta_c$ = +(92.0±8.3)‰, which agrees with the other MM results obtained at R^2 above 0.75. Taking the same subsets of samples, the concomitant ^{13}C signature matches $^{13}\delta_c$ = −(23.3±8.6)‰, indeed at the upper end of the expected LMS $\delta^{13}C(CO)$ variations of −(25−31)‰. Because of that, the MM is likely insensitive to the changes in $\delta^{13}C(CO)$ caused by the contamination (the corresponding R^2 values are below 0.1). Upon the correction using the inferred $^{18}\delta_c$ value, the C1 WAS $\delta^{18}O(CO)$ data agree with B96 (shown with red symbols in Fig. 3). That is, variations in the observed $C^{18}O$ are driven by (i) the seasonal/regional changes in the composition of tropospheric air and by (ii) the degree of mixing or replacement of the latter with the stratospheric

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component that is less variable in ^{18}O. This is seen as stretching of the scattered tropospheric values ([CO] above 60 nmol/mol) towards $\delta^{18}O(CO)$ of around −10‰ at [CO] of 25 nmol/mol, respectively. The corrected C1 $\delta^{13}C(CO)$ data (shown in Fig. 7) are found to be in a ±1‰ agreement with the observations by B96, except for several deep stratospheric samples ([CO] below 40 nmol/mol). The latter were encountered during "ozone hole" conditions and carried extremely low $\delta^{13}C(CO)$ values, which was attributed to the reaction of methane with available free Cl radicals (Brenninkmeijer *et al.*, 1996).

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3.2 Estimate of $\delta^{18}O(O_3)$

[19] The contamination ^{18}O signature inferred here ($^{18}\delta_c$ = +(92.0±8.3)‰) likely pertains to O_3 and is comparable to $\delta^{18}O(O_3)$ values measured in the stratosphere at temperatures about 30 K lower than those encountered in the UT/LMS by C1 (see Table 1 for comparison). If no other factors are involved (see below), this discrepancy in $\delta^{18}O(O_3)$ should be attributed to the local conditions, *i.e.* the higher pressures (typically 240−270 hPa for C1 cruising altitudes) at which O_3 was formed. Indeed, the molecular lifetime (the period through which the species' isotope reservoir becomes entirely renewed, as opposed to the "bulk" lifetime) of O_3 encountered along the C1 flight routes is estimated on the order of minutes to hours at daylight (H. Riede, Max Planck Institute for Chemistry, 2010), thus the isotope composition of the photochemically re-generated O_3 resets quickly according to the local conditions. Virtual absence of sinks, in turn, leads to "freezing" of the $\delta^{18}O(O_3)$ value during night in the UT/LMS. Verifying the current $\delta^{18}O(O_3)$ estimate against the kinetic data, in contrast to the stratospheric cases, is problematic. The laboratory studies on O_3 formation to date have scrutinised the concomitant kinetic isotope effects (KIEs) as a function of temperature at only low pressures (67 mbar); the attenuation of the KIEs with increasing pressure was studied only at room temperatures (see Table 1, also Brenninkmeijer *et al.* (2003) for references). A rather crude attempt may be undertaken by assuming that the formation KIEs become attenuated at higher pressures in a similar (proportional) fashion to that measured at 320 K, however applied to the nominal low-pressure values reckoned at (220−230) K. A decrease in $\delta^{18}O(O_3)$ of about (6−8)‰ is expected from such calculation (*cf.* last row in Table 1), yet accounting for a mere one-half of the (13−15)‰ discrepancy between the stratospheric $\delta^{18}O(O_3)$ values and $^{18}\delta_c$.

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[20] Lower $^{18}\delta_c$ values could result from possible isotope fractionation accompanying the production of the artefact CO. Although not quantifiable here, oxygen KIEs in the $O_3 \rightarrow CO$ conversion chain cannot be ruled out, recalling that the intermediate reaction steps are not identifi-

able and the artefact CO represents at most 4% of all O_3 molecules. Furthermore, the yield λ_{O_3} of CO from O_3 may be lower than unity (see details in Appendix A). On the other hand, the inference that the contamination strength primarily depends on $[O_3]$ indicates that the kinetic fractionation may have greater effect on the carbon isotope ratios of the artefact CO produced (the $^{13}\delta_c$ values) in contrast to the oxygen ones. That is because all reactive oxygen available from O_3 becomes converted to CO, whilst the concomitant carbon atoms are drawn from a virtually unlimited pool whose apparent isotope composition is altered by the magnitude of the ^{13}C KIEs.

[21] Besides KIEs, selectivity in the transfer of O atoms from O_3 to CO affects the resulting $^{18}\delta_c$ value. The terminal O atoms in O_3 are enriched with respect to the molecular (bulk) O_3 composition when the latter is above +70‰ in $\delta^{18}O$ (Janssen, 2005; Bhattacharya *et al.*, 2008), therefore an incorporation of only central O atoms into the artefact CO molecules should result in a reduced apparent $^{18}\delta_c$ value. Such exclusive selection is, however, less likely from the kinetic standpoint and was not observed in available laboratory studies (see Savarino *et al.* (2008) for a review). For instance, Röckmann *et al.* (1998a) established the evidence of direct O transfer from O_3 to the CO produced in alkene ozonolysis. A reanalysis of their results (in light of findings of Bhattacharya *et al.* (2008)) suggests that usually the terminal atoms of the O_3 molecule become transferred (their ratio over the central ones changes from the bulk 2:1 to 1:0 for various species). Considering the alternatives of the O transfer in our case (listed additionally in Table 1), the equiprobable incorporation of the terminal and central O_3 atoms into CO should result in the $\delta^{18}O(O_3)$ value in agreement with the "crude" estimate based on laboratory data given above.

[22] Furthermore, the conditions that supported the reaction of O_3 (or its derivatives) followed by the production of CO are vague. A few hypotheses ought to be scrutinised here. First, a fast $O_3 \rightarrow$ CO conversion must have occurred, owing to short (*i.e.*, fraction of a second) exposure time of the probed air to the contamination. Accounting for the typical C1 air sampling conditions (these are: sampled air pressure of 240−270 hPa and temperature of 220−235 K outboard to 275−300 K inboard, sampling rate of $12.85 \cdot 10^{-3}$ mol s^{-1} corresponding to 350 L STP sampled in 1200 s, inlet/tubing volume gauged to yield exposure times of 0.01 to 0.1 s due to variable air intake rate, $[O_3]$ of 600 nmol/mol), the overall reaction rate coefficient (k_c in Eq. (A3) from Appendix A) must be on the order of $(6 \cdot 10^{-15}/\tau_c)$ molecules^{-1} cm^3, where τ_c is the exposure time. Assuming the case of a gas-phase CO production from a recombining O_3 derivative and an unknown carbonaceous compound X, the reaction rate coefficient for the latter (k in Eq. (A2) in Appendix A) must be unrealistically high, at least $6 \cdot 10^{-10}$ molec^{-1} cm^3 s^{-1} over τ_c = 1/100 s.

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This number decreases proportionally with growing τ_c and [X], if we take less strict exposure conditions. Nonetheless, in order to provide the amounts of artefact CO we detect, a minimum mixing ratio of 20 nmol/mol (or up to 4 µg of C per flight) of X is required, which is not available in the UT/LMS from the species readily undergoing ozonolysis, *e.g.* alkenes.

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[23] Second, a more complex heterogeneous chemistry on the inner surface of the inlet or supplying tubing may be involved. Such can be the tracers' surface adsorption, (catalytic) decomposition of O_3 and its reaction with organics or with surface carbon that also may lead to the production of CO (Oyama, 2000). Evidence exists for the dissociative adsorption of O_3 on the surfaces with subsequent production of the reactive atomic oxygen species (see, *e.g.*, Li *et al.*, 1998, also Oyama, 2000). It is probable that sufficient amounts of organics have remained on the walls of the sampling line exposed to highly polluted tropospheric air, to be later broken down by the products of the heterogeneous decomposition of the ample stratospheric O_3. Unfortunately, the scope for a detailed quantification of intricate surface effects in the C1 CO contamination problem is very limited.

4 Conclusions

[24] Recapitulating, the *in situ* measurements of CO and O_3 allowed us to unambiguously quantify the artefact CO production from O_3 likely in the sample line of the CARIBIC−1 instrumentation. Strong evidence to that is provided by the isotope CO measurements. We demonstrate the ability of the simple mixing model ("Keeling-plot" approach) to single out the contamination isotope signatures even in the case of a large sampling-induced mixing of the air with very different compositions. Obtained as a collateral result, the estimate of the $\delta^{18}O(O_3)$ in the UT/LMS appears adequate, calling, however, for additional laboratory data (*e.g.*, the temperature-driven variations of the O_3 formation KIE at pressures above 100 hPa) for a more unambiguous verification.

Appendix A. Contamination assessment

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We quantify the C1 CO contamination strength (denoted $[CO]_c$, obtained by discriminating the C1 outliers from respective C2 data) in a sequence of regression analyses. We foremost ascertain that no other species or operational parameter (*e.g.* temperature, pressure, flight duration, season, latitude, time of day, *etc.*) measured in C1 appear to determine (*e.g.*, systematically correlate with) $[CO]_c$, except that for $[O_3]$. We hypothesise therefore that a production of artefact

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CO molecules was initiated by O_3 (via either its decomposition or a reaction with an unknown educt) and proceeded with incorporation of carbon (donated by some carbonaceous species X) and oxygen (donated by O_3 or its derivatives) atoms into final CO. Despite that neither the actual reaction chain nor its intermediates are known, it is possible to describe the artefact CO component produced (hereinafter curly brackets {} denote number densities) as

$$\{CO\}_c = \lambda_{O_3} \nu \tau_c , \quad \text{(A1)}$$

where the yield λ_{O_3}, a diagnostic quantity, relates the amount of artefact CO molecules produced to the total number of O_3 molecules consumed in the system, τ_c denotes the reaction time (period throughout which sampled air is exposed to contamination), and ν stands for the overall rate of the reaction chain. The latter, being regarded macroscopically (empirically), is parameterised to account for the order of reaction chain rate with respect to hypothesised reactants (McNaught and Wilkinson, 1997) as

$$\nu = k \{X\}^K \{O_3\}^\kappa , \quad \text{(A2)}$$

where κ and K are the partial orders with respect to X and O_3 number densities, respectively, and k is the rate coefficient. Here it is implied that changes to {X} and {O_3} are negligible throughout the exposure time τ_c (typically < 0.1 s for C1 sample line). As stated above, we find that variations in $\{CO\}_c$ correlate exclusively with variations in {O_3}, hence Eq. (A2) can be reduced by assuming constancy of {X} and K to:

$$\nu_c = k_c \{O_3\}^\kappa . \quad \text{(A3)}$$

Here, $k_c = k\{X\}^K$ (often referred to as pseudo-first-order or "observed" rate coefficient) quantifies the rate of reaction chain exclusively propelled by O_3. Finally, using Eqs. (A1) and (A3), the artefact $\{CO\}_c$ component is expressed as

$$\{CO\}_c = b \cdot \{O_3\}^\kappa , \; b = \lambda_{O_3} k_c \tau_c \quad \text{(A4)}$$

where the constant proportionality factor b integrates the influence of the unknown (and as we explicate below, likely invariable) {X}, k, K and τ_c.

[25] Eq. (A4) defines the regression expression using which we attempt to fit the values of $\{CO\}_c$ as a function of κ, {O_3} and b. In the first regression iteration we keep both κ and b as free parameters, which provides best approximation at κ = 2.06±0.38, suggesting reactions of two O_3 molecules in case elementary reactions constitute the reaction mechanism, or two elementary steps involving O_3 or its derivatives in case a stepwise reaction is involved (McNaught and Wilkinson, 1997). In a subsequent regression iteration we set κ = 2, which yields better (as opposed to the first iteration) estimate of b of (5.19±0.12)·10^{-5} mol/nmol (±1σ,

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adj. R^2 = 0.83, red. χ^2 = 4.0; here the equivalent value in mole fraction units is quoted for the convenience of relating fitted $[CO]_c$ and $[O_3]^2$). At last, we ascertain that the best regression results are obtained particularly at κ = 2, as indicated by the regression statistic (R^2 and χ^2) that asymptotically improves when a set of regressions with neighbouring (*i.e.* below and above 2) integer values of κ is compared. The low uncertainty (within ±3%) associated with the estimate of b confirms an exclusive dependence of the contamination source on the O_3 mixing ratio as well as much similar reaction times τ_c. The regressed value of $[CO]_c$ as a function of $[O_3]$ is presented in Fig. 1 (d) (solid line). It is possible to constrain the overall yield λ_{O_3} of CO molecules in the artefact source chain to be between 0.5 and 1, comparing the magnitude of $[CO]_c$ to the discrepancy between the $[O_3]$ measured in C1 and C2 (±20 nmol/mol, taken equal to the $[O_3]$ bin size owing to the N_2O–O_3 and H_2O–O_3 distributions matching well between the datasets). Lower λ_{O_3} values, otherwise, should have resulted in a noticeable (*i.e.*, greater than 20 nmol/mol) decrease in the C1 O_3 mixing ratios with respect to the C2 levels.

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Appendix B. Corrections to measured $\delta^{13}C(CO)$ values due to the oxygen MIF

[26] Atmospheric O_3 carries an anomalous isotope composition (or mass-independent fractionation, MIF) with a substantially higher relative enrichment in ^{17}O over that in ^{18}O (above +25‰ in $\Delta^{17}O = (\delta^{17}O+1)/(\delta^{18}O+1)^{\beta}-1$, β = 0.528) when compared to the majority of terrestrial oxygen reservoirs that are mass-dependently fractionated (*i.e.*, with $\Delta^{17}O$ of 0‰) (see Brenninkmeijer *et al.* (2003) and refs. therein). CO itself also has an unusual oxygen isotopic composition, possessing a moderate tropospheric MIF of around +5‰ in $\Delta^{17}O(CO)$ induced by the sink KIEs in reaction of CO with OH (Röckmann *et al.*, 1998b; Röckmann *et al.*, 2002) and a minor source effect from the ozonolysis of alkenes (Röckmann *et al.*, 1998a; Gromov *et al.*, 2010). A substantial contamination of CO by O_3 oxygen induces proportional changes to $\Delta^{17}O(CO)$ that largely exceed its natural atmospheric variation. On the other hand, the MIF has implications in the analytical determination of $\delta^{13}C(CO)$, because the presence of $C^{17}O$ species interferes with the mass-spectrometric measurement of the abundances of ^{13}CO possessing the same basic molecular mass (*m/z* is 45). When inferring the exact $C^{17}O/C^{18}O$ ratio in the analysed sample is not possible, analytical techniques usually involve assumptions (*e.g.*, mass-dependently fractionated compositions or a certain non-zero $\Delta^{17}O$ value) with respect to the $C^{17}O$ abundances (Assonov and Brenninkmeijer, 2001). In effect for the C1 CO data, the artefact CO produced

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from O_3 had contributed with unexpectedly high $C^{17}O$ abundances that led to the overestimated $\delta^{13}C(CO)$ analysed. The respective bias ${}^{13}\delta_b$ is quantified using

$$ {}^{13}\delta_b = 7.26 \cdot 10^{-2}\, \Delta^{17}O(CO)\,, \qquad \text{(B1)} $$

where the actual $\Delta^{17}O(CO)$ value is approximated from the natural CO MIF signal ${}^{17}\Delta_n$ and the typical O_3 MIF composition ${}^{17}\Delta_c$ as

$$ \Delta^{17}O(CO) = ({}^{17}\Delta_n\,([CO]_a - [CO]_c) + {}^{17}\Delta_c\,[CO]_c)([CO]_a)^{-1}\,. \qquad \text{(B2)} $$

Here $[CO]_a$ and $[CO]_c$ denote the analysed CO mixing ratio and contamination magnitude, respectively, used in the contamination assessment (see Appendix A, Eq. (A3)) and in calculations with the MM (see Sect. 3.1) . For the purpose of the current estimate it is sufficient to take ${}^{17}\Delta_n$ of +5‰ representing equilibrium enrichments expected in the remote free troposphere and UT/LMS. For the O_3 MIF signature ${}^{17}\Delta_c$, the value of +30‰ (the average $\Delta^{17}O(O_3)$ expected from the kinetic laboratory data at conditions met along the C1 flight routes, see Sect. 3.2 and Table 1) is adopted. The coefficient that proportionates ${}^{13}\delta_b$ and $\Delta^{17}O$ in Eq. (B1) is derived by linearly regressing the $\delta^{13}C(CO)$ biases (simulated using the calculation apparatus detailed by Assonov and Brenninkmeijer, 2001) as a function of $\Delta^{17}O(CO)$ varying within a (0−30)‰ range for the CO with initially unaccounted MIF (*e.g.*, the sample is assumed to be mass-dependently fractionated). It therefore quantifies some extra +(0.726±0.003)‰ in the analysed $\delta^{13}C(CO)$ per every +10‰ of $\Delta^{17}O(CO)$ excess. The most contaminated C1 WAS CO samples at $[O_3]$ above 300 nmol/mol are estimated to bear $\Delta^{17}O(CO)$ of (6−12)‰ corresponding to fractions of (0.10−0.27) of the artefact CO in the sample. Accordingly, the reckoned $\delta^{13}C(CO)$ biases span (0.5−0.9)‰. Although not large, these well exceed the $\delta^{13}C(CO)$ measurement precision of ±0.1‰ and were corrected for, and therefore are taken into account in the calculations with the MM presented in Sect. 3.1.

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Acknowledgements

[27] The authors are indebted to Claus Koeppel, Dieter Scharffe and Dr. Andreas Zahn for their work and expertise on the carbon monoxide and ozone measurements in C1 and C2. Hella Riede is acknowledged for comprehensive estimates of the species lifetimes along the CARIBIC flight routes. We are grateful to Dr. Taku Umezawa, Dr. Angela K. Baker, Dr. Emma C. Leedham, Dr. Sergey Assonov, the anonymous reviewer and Dr. Jan Kaiser for the helpful discussions and comments on the manuscript.

References

Assonov, S. S. and Brenninkmeijer, C. A. M.: A new method to determine the ^{17}O isotopic abundance in CO_2 using oxygen isotope exchange with a solid oxide, *Rapid Commun. Mass Spectrom.*, **15**, 2426−2437, doi: 10.1002/rcm.529, 2001.

Assonov, S. S. and Brenninkmeijer, C. A. M.: A redetermination of absolute values for $^{17}R_{VPDB\text{-}CO2}$ and $^{17}R_{VSMOW}$, *Rapid Commun. Mass Spectrom.*, **17**, 1017−1029, doi: 10.1002/Rcm.1011, 2003.

Assonov, S. S., Brenninkmeijer, C. A. M., Koeppel, C., and Röckmann, T.: CO_2 isotope analyses using large air samples collected on intercontinental flights by the CARIBIC Boeing 767, *Rapid Commun. Mass Spectrom.*, **23**, 822−830, doi: 10.1002/rcm.3946, 2009.

Bhattacharya, S. K., Pandey, A., and Savarino, J.: Determination of intramolecular isotope distribution of ozone by oxidation reaction with silver metal, *J. Geophys. Res. Atm.*, **113**, D03303, doi: 10.1029/2006jd008309, 2008.

Brenninkmeijer, C. A. M.: Measurement of the abundance of ^{14}CO in the atmosphere and the $^{13}C/^{12}C$ and $^{18}O/^{16}O$ ratio of atmospheric CO with applications in New Zealand and Antarctica, *J. Geophys. Res. Atm.*, **98**, 10595−10614, doi: 10.1029/93JD00587, 1993.

Brenninkmeijer, C. A. M., Müller, R., Crutzen, P. J., Lowe, D. C., Manning, M. R., Sparks, R. J., and van Velthoven, P. F. J.: A large ^{13}CO deficit in the lower Antarctic stratosphere due to "Ozone Hole" Chemistry: Part I, Observations, *Geophys. Res. Lett.*, **23**, 2125−2128, doi: 10.1029/96gl01471, 1996.

Brenninkmeijer, C. A. M. and Röckmann, T.: Principal factors determining the $^{18}O/^{16}O$ ratio of atmospheric CO as derived from observations in the southern hemispheric troposphere and lowermost stratosphere, *J. Geophys. Res. Atm.*, **102**, 25477−25485, doi: 10.1029/97JD02291, 1997.

Brenninkmeijer, C. A. M., Crutzen, P. J., Fischer, H., Gusten, H., Hans, W., Heinrich, G., Heintzenberg, J., Hermann, M., Immelmann, T., Kersting, D., Maiss, M., Nolle, M., Pitscheider, A., Pohlkamp, H., Scharffe, D., Specht, K., and Wiedensohler, A.: CARIBIC – Civil aircraft for global measurement of trace gases and aerosols in the tropopause region, *J. Atmos. Oceanic Technol.*, **16**, 1373−1383, doi: 10.1175/1520-0426(1999)016<1373:Ccafgm>2.0.Co;2, 1999.

Brenninkmeijer, C. A. M., Koeppel, C., Röckmann, T., Scharffe, D. S., Bräunlich, M., and Gros, V.: Absolute measurement of the abundance of atmospheric carbon monoxide, *J. Geophys. Res. Atm.*, **106**, 10003−10010, doi: 10.1029/2000jd900342, 2001.

Brenninkmeijer, C. A. M., Janssen, C., Kaiser, J., Röckmann, T., Rhee, T. S., and Assonov, S. S.: Isotope effects in the chemistry of atmospheric trace compounds, *Chem. Rev.*, **103**, 5125−5161, doi: 10.1021/Cr020644k, 2003.

Brenninkmeijer, C. A. M., Crutzen, P., Boumard, F., Dauer, T., Dix, B., Ebinghaus, R., Filippi, D., Fischer, H., Franke, H., Frieß, U., Heintzenberg, J., Helleis, F., Hermann, M., Kock, H. H., Koeppel, C., Lelieveld, J., Leuenberger, M., Martinsson, B. G., Miemczyk, S., Moret, H. P., Nguyen, H. N., Nyfeler, P., Oram, D., O'Sullivan, D., Penkett, S., Platt, U., Pupek, M., Ramonet, M., Randa, B., Reichelt, M., Rhee, T. S., Rohwer, J., Rosenfeld, K., Scharffe, D., Schlager, H., Schumann, U., Slemr, F., Sprung, D., Stock, P., Thaler, R., Valentino, F., van Velthoven, P., Waibel, A., Wandel, A., Waschitschek, K., Wiedensohler, A., Xueref-Remy, I., Zahn, A., Zech, U., and Ziereis, H.: Civil Aircraft for the regular investigation of the atmosphere based on an instrumented container: The new CARIBIC system, *Atmos. Chem. Phys.*, **7**, 4953−4976, doi: 10.5194/acp-7-4953-2007, 2007.

Coplen, T. B.: Reporting of stable hydrogen, carbon, and oxygen isotopic abundances (Technical Report), *Pure Appl. Chem.*, **66**, 273−276, doi: 10.1351/pac199466020273, 1994.

Craig, H.: Isotopic standards for carbon and oxygen and correction factors for mass-spectrometric analysis of carbon dioxide, *Geochim. Cosmochim. Acta*, **12**, 133−149, doi: 10.1016/0016-7037(57)90024-8, 1957.

Gonfiantini, R.: Standards for Stable Isotope Measurements in Natural Compounds, *Nature*, **271**, 534−536, 1978.

Gromov, S., Jöckel, P., Sander, R., and Brenninkmeijer, C. A. M.: A kinetic chemistry tagging technique and its application to modelling the stable isotopic composition of atmospheric trace gases, *Geosci. Model Dev.*, **3**, 337−364, doi: 10.5194/gmd-3-337-2010, 2010.

Guenther, J., Erbacher, B., Krankowsky, D., and Mauersberger, K.: Pressure dependence of two relative ozone formation rate coefficients, *Chem. Phys. Lett.*, **306**, 209−213, doi: 10.1016/S0009-2614(99)00469-8, 1999.

Janssen, C., Guenther, J., Krankowsky, D., and Mauersberger, K.: Temperature dependence of ozone rate coefficients and isotopologue fractionation in $^{16}O-^{18}O$ oxygen mixtures, *Chem. Phys. Lett.*, **367**, 34−38, doi: 10.1016/S0009-2614(02)01665-2, 2003.

Janssen, C.: Intramolecular isotope distribution in heavy ozone ($^{16}O^{18}O^{16}O$ and $^{16}O^{16}O^{18}O$), *J. Geophys. Res. Atm.*, **110**, D08308, doi: 10.1029/2004jd005479, 2005.

Johnston, J. C. and Thiemens, M. H.: The isotopic composition of tropospheric ozone in three environments, *J. Geophys. Res. Atm.*, **102**, 25395−25404, doi: 10.1029/97jd02075, 1997.

Krankowsky, D., Bartecki, F., Klees, G. G., Mauersberger, K., Schellenbach, K., and Stehr, J.: Measurement of heavy isotope enrichment in tropospheric ozone, *Geophys. Res. Lett.*, **22**, 1713−1716, doi: 10.1029/95gl01436, 1995.

Krankowsky, D., Lämmerzahl, P., Mauersberger, K., Janssen, C., Tuzson, B., and Röckmann, T.: Stratospheric ozone isotope fractionations derived from collected samples, *J. Geophys. Res. Atm.*, **112**, D08301, doi: 10.1029/2006jd007855, 2007.

Li, W., Gibbs, G. V., and Oyama, S. T.: Mechanism of Ozone Decomposition on a Manganese Oxide Catalyst. 1. In Situ Raman Spectroscopy and Ab Initio Molecular Orbital Calculations, *J. Am. Chem. Soc.*, **120**, 9041−9046, doi: 10.1021/ja981441+, 1998.

Mauersberger, K.: Measurement of Heavy Ozone in the Stratosphere, *Geophys. Res. Lett.*, **8**, 935−937, doi: 10.1029/Gl008i008p00935, 1981.

McNaught, A. D. and Wilkinson, A.: IUPAC. Compendium of Chemical Terminology (the "Gold Book"), XML on-line corrected version: http://goldbook.iupac.org (2006-) created by M. Nic, J. Jirat, B. Kosata; updates compiled by A. Jenkins, doi: 10.1351/goldbook.O04322, 1997.

Natrella, M.: NIST/SEMATECH e-Handbook of Statistical Methodsed., edited by: Croarkin, C. and Tobias, P., NIST/SEMATECH, http://www.itl.nist.gov/div898/handbook/ (last access: 07 May 2014), 2003.

Novelli, P. C., Masarie, K. A., and Lang, P. M.: Distributions and recent changes of carbon monoxide in the lower troposphere, *J. Geophys. Res.*, **103**, 19015−19033, doi: 10.1029/98jd01366, 1998.

Novelli, P. C., Masarie, K. A., Lang, P. M., Hall, B. D., Myers, R. C., and Elkins, J. W.: Reanalysis of tropospheric CO trends: Effects of the 1997−1998 wildfires, *J. Geophys. Res.*, **108**, 4464, doi: 10.1029/2002jd003031, 2003.

Oyama, S. T.: Chemical and Catalytic Properties of Ozone, *Catal. Rev. Sci. Eng.*, **42**, 279−322, doi: 10.1081/cr-100100263, 2000.

Pan, L. L., Randel, W. J., Gary, B. L., Mahoney, M. J., and Hintsa, E. J.: Definitions and sharpness of the extratropical tropopause: A trace gas perspective, *J. Geophys. Res. Atm.*, **109**, D23103, doi: 10.1029/2004jd004982, 2004.

Röckmann, T., Brenninkmeijer, C. A. M., Neeb, P., and Crutzen, P. J.: Ozonolysis of nonmethane hydrocarbons as a source of the observed mass independent oxygen isotope enrichment in tropospheric CO, *J. Geophys. Res. Atm.*, **103**, 1463−1470, doi: 10.1029/97JD02929, 1998a.

Röckmann, T., Brenninkmeijer, C. A. M., Saueressig, G., Bergamaschi, P., Crowley, J. N., Fischer, H., and Crutzen, P. J.: Mass-independent oxygen isotope fractionation in atmospheric CO as a result of the reaction CO+OH, *Science*, **281**, 544−546, doi: 10.1126/science.281.5376.544, 1998b.

Röckmann, T., Jöckel, P., Gros, V., Bräunlich, M., Possnert, G., and Brenninkmeijer, C. A. M.: Using ^{14}C, ^{13}C, ^{18}O and ^{17}O isotopic variations to provide insights into the high northern latitude surface CO inventory, *Atmos. Chem. Phys.*, **2**, 147−159, doi: 10.5194/acp-2-147-2002, 2002.

Savarino, J., Bhattacharya, S. K., Morin, S., Baroni, M., and Doussin, J. F.: The NO+O_3 reaction: A triple oxygen isotope perspective on the reaction dynamics and atmospheric implications for the transfer of the ozone isotope anomaly, *J. Chem. Phys.*, **128**, 194303, doi: 10.1063/1.2917581, 2008.

Savarino, J. and Morin, S.: The N, O, S Isotopes of Oxy-Anions in Ice Cores and Polar Environments, in: Handbook of Environmental Isotope Geochemistry, edited by: Baskaran, M., Advances in Isotope Geochemistry, Springer Berlin Heidelberg, 835−864, 2012.

Scharffe, D., Slemr, F., Brenninkmeijer, C. A. M., and Zahn, A.: Carbon monoxide measurements onboard the CARIBIC passenger aircraft using UV resonance fluorescence, *Atmos. Meas. Tech.*, **5**, 1753−1760, doi: 10.5194/amt-5-1753-2012, 2012.

Schinke, R., Grebenshchikov, S. Y., Ivanov, M. V., and Fleurat-Lessard, P.: Dynamical Studies Of The Ozone Isotope Effect: A Status Report, *Annu. Rev. Phys. Chem.*, **57**, 625−661, doi: 10.1146/annurev.physchem.57.032905.104542, 2006.

Stevens, C. M., Kaplan, L., Gorse, R., Durkee, S., Compton, M., Cohen, S., and Bielling, K.: The Kinetic Isotope Effect for Carbon and Oxygen in the Reaction CO+OH, *Int. J. Chem. Kinet.*, **12**, 935−948, doi: 10.1002/kin.550121205, 1980.

Vicars, W. C., Bhattacharya, S. K., Erbland, J., and Savarino, J.: Measurement of the ^{17}O-excess ($\Delta^{17}O$) of tropospheric ozone using a nitrite-coated filter, *Rapid Commun. Mass Spectrom.*, **26**, 1219−1231, doi: 10.1002/rcm.6218, 2012.

Vicars, W. C. and Savarino, J.: Quantitative constraints on the ^{17}O-excess ($\Delta^{17}O$) signature of surface ozone: Ambient measurements from 50°N to 50°S using the nitrite-coated filter technique, *Geochim. Cosmochim. Acta*, **135**, 270−287, doi: 10.1016/j.gca.2014.03.023, 2014.

Zahn, A., Brenninkmeijer, C. A. M., Maiss, M., Scharffe, D. H., Crutzen, P. J., Hermann, M., Heintzenberg, J., Wiedensohler, A., Güsten, H., Heinrich, G., Fischer, H., Cuijpers, J. W. M., and van Velthoven, P. F. J.: Identification of extratropical two-way troposphere-stratosphere mixing based on CARIBIC measurements of O_3, CO, and ultrafine particles, *J. Geophys. Res.*, **105**, 1527−1535, doi: 10.1029/1999jd900759, 2000.

Zahn, A., Brenninkmeijer, C. A. M., Asman, W. A. H., Crutzen, P. J., Heinrich, G., Fischer, H., Cuijpers, J. W. M., and van Velthoven, P. F. J.: Budgets of O_3 and CO in the upper troposphere:

CARIBIC passenger aircraft results 1997–2001, *J. Geophys. Res. Atm.*, **107**, 4337, doi: 10.1029/2001jd001529, 2002.

Zahn, A., Weppner, J., Widmann, H., Schlote-Holubek, K., Burger, B., Kühner, T., and Franke, H.: A fast and precise chemiluminescence ozone detector for eddy flux and airborne application, *Atmos. Meas. Tech.*, **5**, 363−375, doi: 10.5194/amt-5-363-2012, 2012.

Tables

Table 1. Ozone $^{18}O/^{16}O$ isotope ratios from literature and this study

Domain	T (K)	P (hPa)	$\delta^{18}O(O_3)$ (‰)	Remarks
Stratosphere	190−210	13−50	83−93 (<3)	[1]
UT/LMS	220−235	240−270	89−95 (8)	[2]
			84−88 (6)	T
			91−98 (9)	TC
			112−124 (17)	C
Laboratory	190−210	67	87−97 (6)	[3]
	220−235	67	102−110 (6)	[3]
	220−235	240−270	95−103	[4]

Notes: Values in parentheses denote the average of the estimates' standard errors. The expected O_3 isotope composition on the VSMOW scale is calculated from the O_3 enrichments reported relative to O_2 using $\delta^{18}O(O_3)_{VSMOW} = \delta^{18}O(O_2)_{VSMOW} + {}^{18}\delta(O_3)_{Air\text{-}O_2} + [\delta^{18}O(O_2)_{VSMOW} \times {}^{18}\varepsilon(O_3)_{Air\text{-}O_2}]$.

[1] Observations (see Krankowsky *et al.* (2007) and refs. therein), lowermost values (19−25 km). Quoted temperature range is derived by matching measured $\delta^{18}O(O_3)$ and laboratory data (see note [3]).

[2] This study, C1 observations (10−12 km). Letters denote the estimates derived using the data from Bhattacharya *et al.* (2008) and assuming only terminal (T), only central (C) and equiprobable terminal and central (TC) O_3 atoms transfer to the artefact CO.

[3] Calculated using the laboratory KIE temperature dependence data summarised by Janssen *et al.* (2003).

[4] Calculated assuming a pressure dependence of the O_3 formation KIE similar to that measured at 320 K (see Guenther *et al.* (1999) and refs. therein).

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836 **Figures**

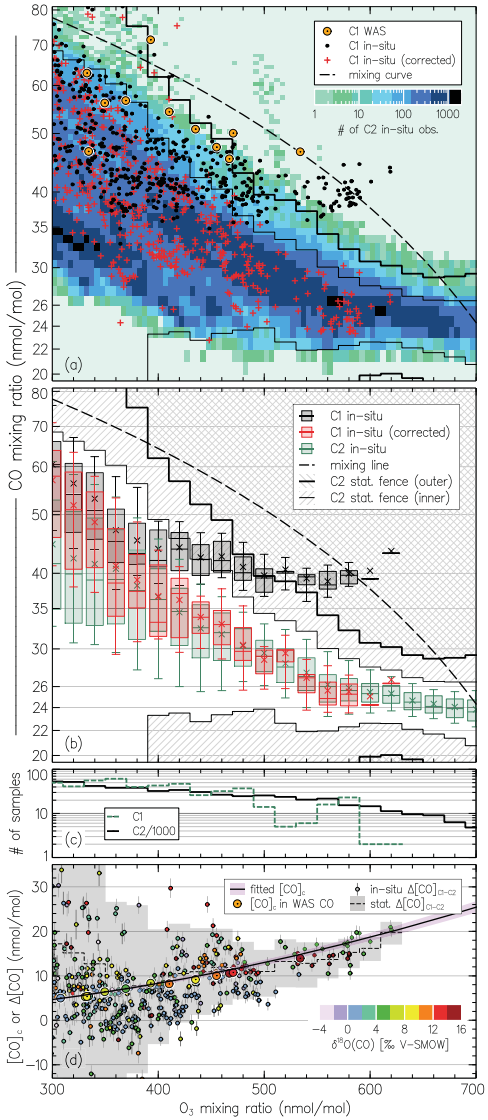
C1 WAS
C1 in-situ
C1 in-situ (corrected)
mixing curve
1
10
100
1000
# of C2 in-situ obs.
(a)
CO mixing ratio (nmol/mol)
C1 in-situ
C1 in-situ (corrected)
C2 in-situ
mixing line
C2 stat. fence (outer)
C2 stat. fence (inner)
(b)
# of samples
C1
C2/1000
(c)
[CO]c or Δ[CO] (nmol/mol)
fitted [CO]c
[CO]c in WAS CO
in-situ Δ[CO]C1–C2
stat. Δ[CO]C1–C2
−4
0
4
8
12
16
δ18O(CO) [‰ V-SMOW]
(d)
300
400
500
600
700
O3 mixing ratio (nmol/mol)

Fig. 1. (a) Distribution of CO mixing ratios as a function of concomitant O_3 mixing ratios measured by CARIBIC in the LMS ([O_3]>300 nmol/mol). The shaded area is the two-dimensional histogram of the C2 measurements (all C2 data obtained until June 2013) counted in 5×1 nmol/mol size [O_3]×[CO] bins, thus darker areas emphasise greater numbers of particular CO–O_3 pairs observed. Small symbols denote the original C1 *in situ* measurements (black) and corrected for the artefacts (red); the C1 WAS analyses (11 of total 408) are shown with large symbols. Thin and thick step-lines demark the inner and outer statistical fences (ranges outside which the data points are considered mild or extreme outliers, see text) of the C2 data, respectively. The dashed curve exemplifies compositions expected from the linear mixing of very different (*e.g.*, tropospheric and stratospheric) end-members. (b) Statistics on CO mixing ratios from C1 and C2 data shown in box-and-whisker diagrams for samples clustered in 20 nmol/mol O_3 bins (whiskers represent $9^{th}/91^{st}$ percentiles). (c) Sample statistic for each CARIBIC dataset (note the C2 figures scaled down by a factor of 1000). (d) Estimates of the C1 *in situ* CO contamination strength $[CO]_c$ as a function of [O_3] (solid line) obtained by fitting the difference Δ[CO] between the C2 and C1 *in situ* [CO] (small symbols) as detailed in Appendix A (Eq. (A1)). Step line shows the Δ[CO] for the statistical averages (the shaded area equals the height of the inner statistical fences of the C2 data). Large symbols denote the estimates of C_c in the C1 WAS data (slight variations *vs.* the *in situ* data are due to the sample mixing effects, see Sect. 3). Colour denotes the respective C1 WAS $\delta^{18}O(CO)$ (note that typically 6–7 *in situ* measurements correspond to one WAS sample).

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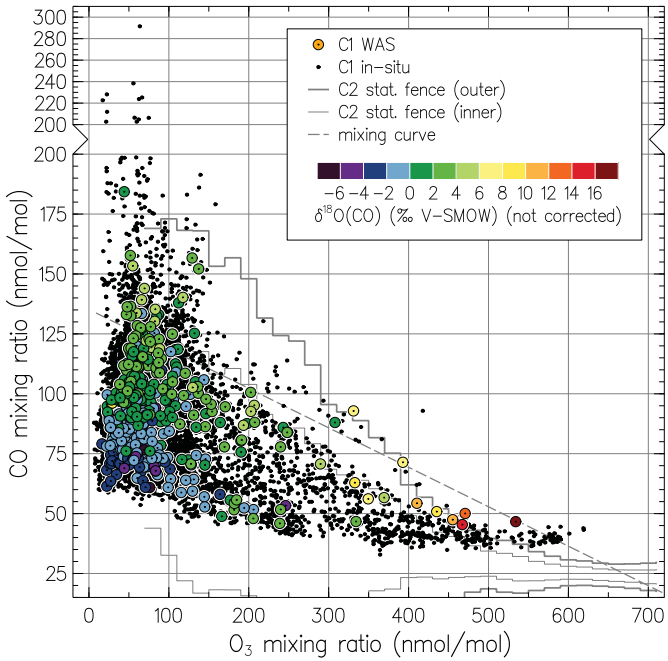


Fig. 2. (accompanies Fig. 1) Carbon monoxide and ozone mixing ratios measured in C1. Small black symbols denote the C1 *in situ* measurements (n = 12753). The C1 WAS analyses (n = 408) are shown with large symbols; colour denotes the concomitant $\delta^{18}O(CO)$ measurements. Thin and thick step-lines denote the inner and outer statistical fences of the C2 data, respectively. The dashed curve exemplifies compositions expected from the linear mixing of tropospheric and stratospheric end-members (see caption to Fig. 1 for details).

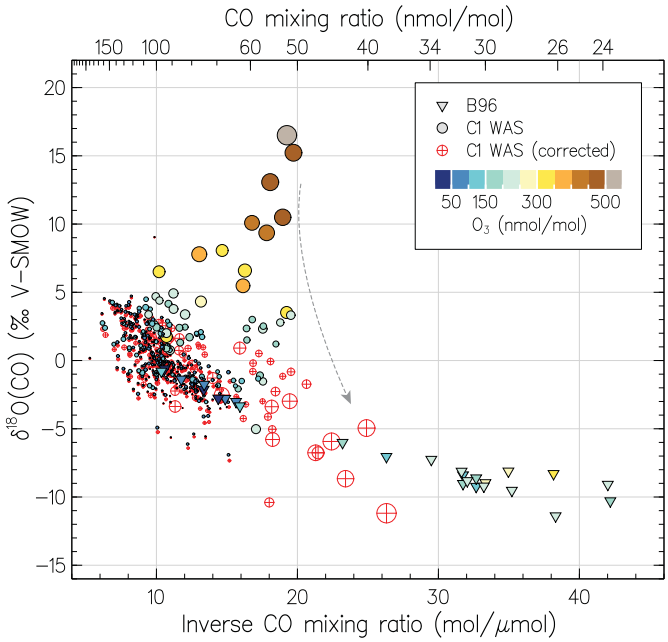
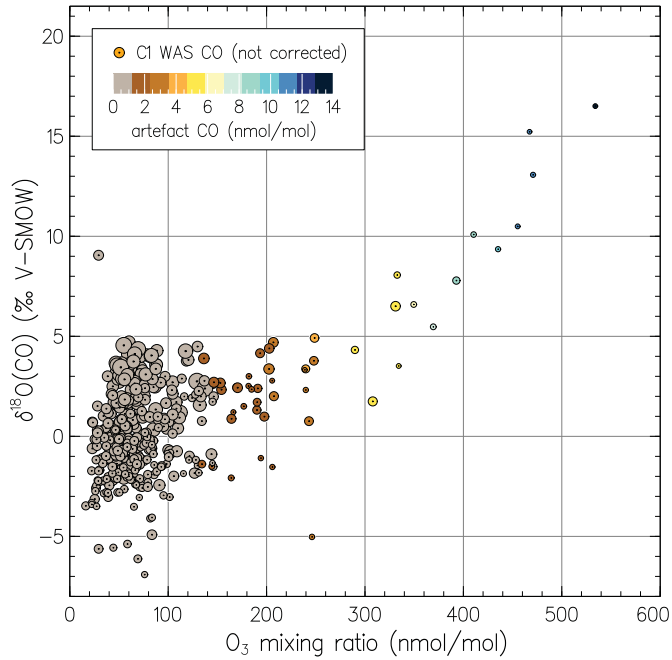


Fig. 3. $^{18}O/^{16}O$ isotope composition of CO as a function of its reciprocal mixing ratio. Triangles present the data from the remote SH UT/LMS obtained by Brenninkmeijer *et al.* (1996) (B96). Colour refers to the concomitantly observed O_3 abundances; note the extremely low [O_3] encountered by B96 in the Antarctic "ozone hole" conditions. Filled and hollow circles denote the original and corrected (as exemplified by the dashed arrow) C1 WAS data, respectively, with the symbol size scaling proportional to the estimated contamination magnitude (see text).



869 Fig. 4. Measured C1 WAS $\delta^{18}O(CO)$ (not corrected for artefacts) as a function of concomitant O_3 mixing
870 ratio. Symbol colour denotes the artefact CO component (integral $[CO]_c$ per each WAS); symbol size
871 scales proportionally to the WAS CO mixing ratio corrected for artefacts (see Sect. 3 for details).

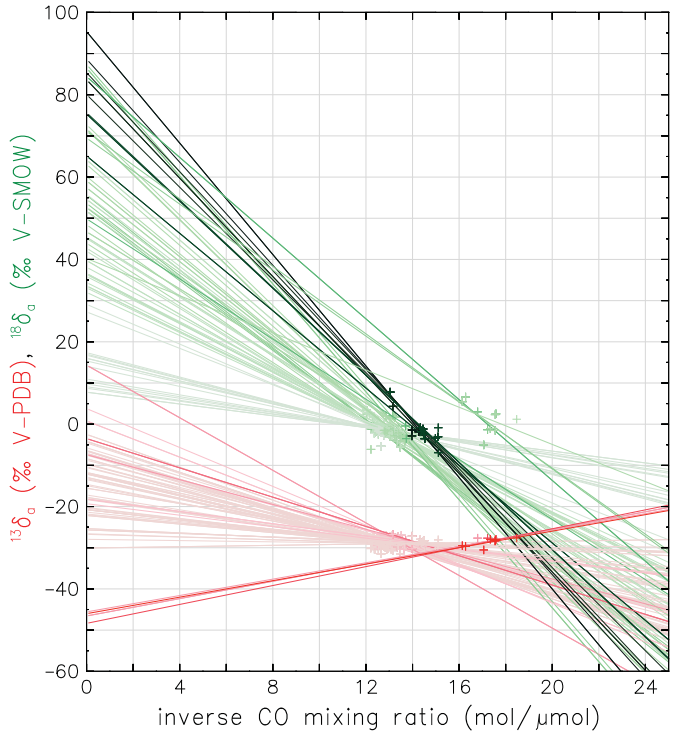


Fig. 5. Keeling plot of the data used in the calculations with the mixing model (MM). The C1 WAS isotope CO measurements are shown with symbols, solid lines denote the linear regressions through the various sets of samples selected by the MM (n = 80 sets are plotted). Colours refer to the $\delta^{13}C$ (red) and $\delta^{18}O$ (green) data, colour intensity indicates the coefficient of determination (R^2) of each regression, respectively. Darker colours denote higher R^2 values, with maxima of 0.92 for $\delta^{18}O$ and 0.54 for $\delta^{13}C$ data, respectively. The inferred contamination signatures ($^i\delta_c$) are found at $([CO]_a)^{-1} \rightarrow 0$. Regression uncertainties are shown in Fig. 6. Note that because different subsets of samples contain same data points, some of the symbols are plotted over (*i.e.*, not all symbols contributing to a particular regression case may be seen).

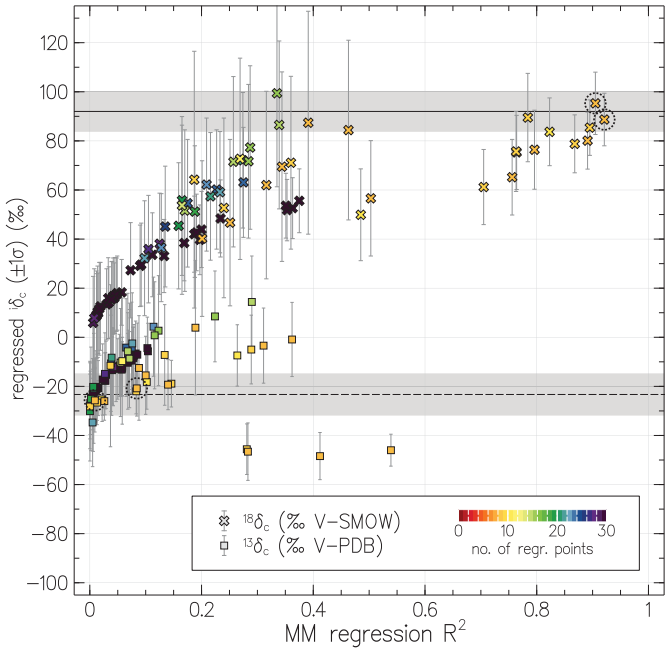


Fig. 6. Results of the regression calculation with the MM. Shown with symbols are the contamination source isotope signatures $^{i}\delta_c$ as a function of the respective coefficient of determination (R^2). Colour denotes the number of samples in each subset selected. Solid and dashed lines present the best guess ±1 standard deviation of the mean for the $^{18}\delta_c$ and $^{13}\delta_c$ estimates. Dashed circles mark the estimates obtained at highest R^2 for $^{18}\delta_c$ regression (above 0.9). See text for details.

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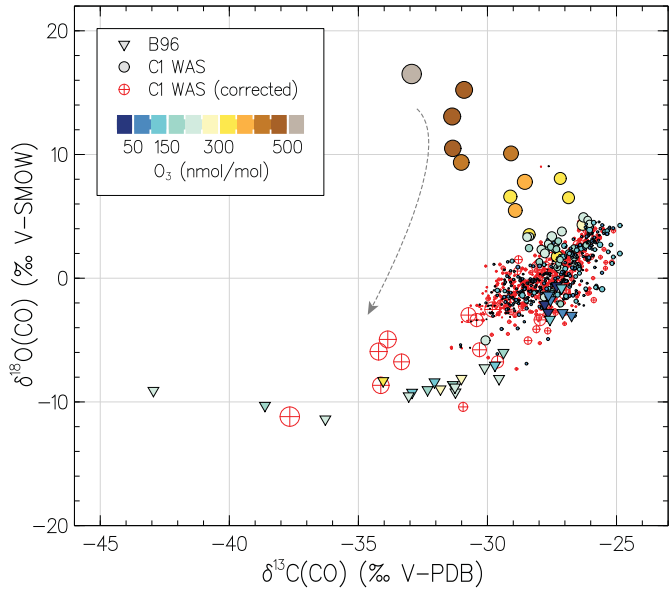

Fig. 7. $^{18}O/^{16}O$ and $^{13}C/^{12}C$ isotope composition of CO measured in C1. Triangles present the data from the remote SH UT/LMS obtained by Brenninkmeijer *et al.* (1996) (B96). Colour refers to the concomitantly observed O_3 abundances; note the extremely low [O_3] encountered by B96 in the Antarctic ozone-hole conditions. Filled and hollow circles denote the original and corrected (as exemplified by the dashed arrow) C1 WAS data, respectively, with the symbol size scaling proportional to the estimated contamination magnitude (see text for details).

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s K and κ describes the stoichiometry of the system with respect to ozone, *i.e.* how many (rate-determining) reactions of X and O_3 ((with the individual unknown rate coefficients ${}^{X}k_r$ and ${}^{O_3}k_r$or its derivatives)) may have lead to production of one artefact CO, respectively. The yield λ_{O_3}, a diagnostic quantity, relates the amount of artefact CO molecules produced to the total number of O_3 molecules consumed in the system. Based on Eq. (A1), the functional dependence of the artefact CO component (denoted $[CO]_c$, obtained by discriminating the C1 outliers from respective C2 data) on $[O_3]$ or [X] is generally formulated as (abundances in number density units are used)

$$C_c = \int_{\tau_c} \prod_{\kappa} {}^{O_3}k_r [O_3] \prod_{K} {}^{X}k_r [X] dt \ , \qquad \text{(A2)}$$

where τ_c denotes the contamination reaction time.

[1]

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, K and [X] (the latter was chosen iteratively from a set of carbonaceous species measured). Practically, however, this regression analysis ascertains that variations in $[CO]_c$ are exhaustively described using $[O_3]$ and κ. Furthermore, we find that no other species or operational parameter (e.g. temperature, pressure, flight duration, latitude, etc.) measured in C1 appear to determine (correlate with) $[CO]_c$. Based on this, we can reduce Eq. (A2) to its final, simpler form, viz.

$$CO]_c = \lambda_{O_3} k_c [O_3]^{\kappa} \tau_c \ , \quad \text{(A3)}$$

where k_c denotes the overall pseudo-first-order rate coefficient of the reaction chain that is exclusively propelled by O_3. The product $(\lambda_{O_3} k_c \tau_c)$ thus integrates the influence of the unknown (and likely invariable) [X], ${}^{X}k_r$, K and τ_c

[2]

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the product $(\lambda_{O_3} k_c \tau_c)$ that proportionates the CO contamination strength and $[O_3]$ is found to be

[3]