

Relating hygroscopicity and optical properties to chemical composition and structure of secondary organic aerosol particles generated from the ozonolysis of α -pinene

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Abstract

Secondary Organic Aerosols (SOA) were generated from the ozonolysis of α -pinene in the CESAM simulation chamber. The SOA formation and aging were studied by following their optical, hygroscopic and chemical properties. The optical properties were investigated by determining the particle Complex Refractive Index (CRI). The hygroscopicity was quantified by measuring the effect of RH on the particle size (size growth factor, GF) and on the scattering coefficient (scattering growth factor, $f(RH)$). The oxygen to carbon atomic ratios (O:C) of the particle surface and bulk were used as a sensitive parameter to correlate the changes in

hygroscopic and optical properties of the SOA composition during their formation and aging in CESAM.

The real CRI at 525 nm wavelength decreased from 1.43-1.60 (± 0.02) to 1.32-1.38 (± 0.02) during the SOA formation. The decrease in the real CRI correlated to the O:C decrease from 0.68 (± 0.20) to 0.55 (± 0.16). In contrast, the GF remained roughly constant over the reaction time, with values of 1.02-1.07 (± 0.02) at 90% ($\pm 4.2\%$) RH. Simultaneous measurements of O:C of the particle surface revealed that the SOA was not composed of a homogeneous mixture, but contained less oxidised species at the surface which may limit water absorption. In addition, an apparent change in both mobility diameter and scattering coefficient with increasing RH from 0 to 30% was observed for SOA after 14 hours of reaction. We postulate that this change could be due to a change in the viscosity of the SOA from a predominantly glassy state to a predominantly liquid state.

1 Introduction

Organic compounds are known to account for a large fraction of atmospheric aerosol, ranging between 20 and 90% of the total particle mass (Kanakidou et al., 2005). In particular, secondary organic aerosol (SOA), formed by the condensation of gas-phase oxidation products of volatile organic compounds (VOCs), is a major constituent of the total organic aerosol (Turpin and Huntzicker, 1995; Zhang et al., 2007).

SOA size ranges between ten to hundreds of nanometers. Particles in this size range have long atmospheric lifetimes and scatter solar radiation. SOA can also change clouds properties by acting as Cloud Condensation Nuclei (CCN) (Saxena et al., 1995; Lohmann and Feichter, 2005; Novakov and Penner, 1993; Rivera-Carpio et al., 1996; Matsumoto et al., 1997).

There are large uncertainties in estimating the impact of SOA on climate due to their complexity and the limited range of measurements available (Kanakidou et al., 2005). SOA precursors produce a large number of oxidation products (Goldstein and Galbally, 2009), resulting in many possible chemical reaction pathways (de Gouw et al., 2005; Hallquist et al., 2009). In addition, during their lifetime in the atmosphere, SOA may undergo several physical and chemical aging processes altering their chemical composition (Kalberer et al., 2004; Baltensperger, 2005;

Yasmeen et al., 2012) and size distribution (Andreae, 2009). As a result, atmospheric SOA contain many organic compounds with a large variety of structures, chain lengths, functionalities and degrees of oxidation (Kroll and Seinfeld, 2008; Jimenez et al., 2009). Therefore SOA possess a wide range of hygroscopic and optical properties (Lambe et al., 2013; Suda et al., 2012).

In global climate models, the direct radiative effect of SOA is currently described by adopting a constant Complex Refractive Index (CRI) and a single size Growth Factor (GF). Depending on the model, the adopted real part of the CRI at visible wavelengths ranges from 1.45 to 1.60 (Kinne et al., 2003; Pere et al., 2011; Zaveri et al., 2010). Some models assume that SOA absorb weakly solar radiation, and set the imaginary part of the CRI near 0.006, while others ignore the absorption by SOA. Concerning hygroscopic properties, a single size GF of SOA derived from limited available data is used (O'Donnell et al., 2011; Hoyle et al., 2009). However, field measurements shows that the hygroscopic and optical properties of SOA are not static and depend on their origin and transport in the atmosphere (Duplissy et al., 2010; Jimenez et al., 2009; Chang et al., 2010; Dinar et al., 2008). As a consequence, laboratory investigations have started to explore the change of SOA properties during their lifetime. The ozonolysis of α -pinene ozonolysis is one of the most well studied SOA systems (α -pinene-O₃ SOA), as α -pinene is a significant biogenic VOC in many regions, and its ozonolysis plays an important role in SOA formation (Hallquist et al., 2009; Griffin et al., 1999; Yu et al., 1999; Kavouras et al., 1998, 1999). α -pinene-O₃ SOA is also generally considered as a model for many biogenic VOCs containing an endocyclic double bond (Guenther et al., 1995). Saathoff et al. (2003) conducted experiments of α -pinene-O₃ SOA in a simulation chamber and observed an increase of the GF at 90% RH from 1.080 to 1.106. Cocker et al. (2001) reported also an increase of the GF at 85% RH from 1.065 to 1.11 within 6 hours. In contrast, Warren et al. (2009) and Qi et al. (2010) reported a constant GF at 90% RH for α -pinene-O₃ SOA over 6 hours of reaction. Finally, an increase of the real CRI from 1.39 to 1.52 at $\lambda = 532$ nm during the formation of the α -pinene-O₃ SOA has been reported by Kim et al. (2013). To date, none of the previous studies have simultaneously determined the hygroscopic and optical properties, and their evolution with the particle chemical composition during SOA formation and ageing.

To accurately quantify the SOA impacts on climate, it is critical to determine the hygroscopic behaviour of both the size distribution and optical properties as well as the dependence of these

properties on the chemical composition. Atmospheric simulation chambers are powerful tools to study the physical, chemical, optical and hygroscopic properties of SOA and follow their changes along their lifecycle by simulating their formation and aging due to oxidation and photolysis in the atmosphere (Meyer *et al.*, 2008; Henry and Donahue, 2012; Tritscher *et al.*, 2011). In this work, we take advantage of the long aerosol lifetime in the CESAM chamber (French acronym for Experimental Multiphasic Atmospheric Simulation Chamber, Wang *et al.*, 2011) to set-up formation and aging experiments of α -pinene-O₃ SOA in order to characterise the evolution of both the SOA optical properties and hygroscopicity. The objective is to examine the evolution of these physical properties and to relate them to the aerosol chemical composition.

2 Methods

Measurements have been conducted in the humidity-controlled simulation chamber CESAM (Wang *et al.*, 2011), which permits the study of the formation and aging of SOA over long periods of time, and under various relevant atmospheric conditions (temperature, relative humidity, pressure, gas phase concentration, etc.).

Experiments were conducted to simultaneously measure different parameters:

- CRI at $\lambda=525\text{nm}$, a specific wavelength in the mid-visible; the Complex Refractive Index (CRI, $m = n - ik$) is an important parameter to link the physical and chemical properties of the SOA and its ability to interact with radiation, allowing a description of the scattering and absorbing characteristics of SOA.
- GF(RH), size growth factor, the ratio of the particle diameter at a given RH to the particle diameter at low RH, for one selected size of particles; this parameter is used to characterise the hygroscopic properties of the SOA and thus, it is an indicator of its ability to act as CCN.
- $f(\text{RH})$, scattering growth factor, the ratio of the scattering coefficient (σ_{scat}) at high RH to the σ_{scat} at low RH, for the entire size distribution; this parameter allows the study of the effect of water absorption on the scattering properties of SOA.
- O:C, the oxygen-to-carbon ratio, for both the bulk and the surface composition of the SOA; it is well known that particle composition, in particular that of the surface, strongly influences the water uptake ability of the particle and its CCN potential (McFiggans *et al.*, 2006; Dusek *et al.*, 2006; Hatch *et al.*, 2008; Moussa *et al.*, 2009; McIntire *et al.*, 2010; Semeniuk *et al.*, 2007; Lamb *et al.*, 2011; Wong *et al.*, 2011; Mei *et al.*, 2013; Rickards *et al.*, 2013). To encompass the difficulty of representing the full molecular composition,

the O:C of SOA has been included in global climate models to provide a description of the aerosol aging (Tost and Pringle, 2012). Furthermore, a number of recent studies have reported a positive correlation between hygroscopicity and bulk O:C for both laboratory and ambient SOA (Massoli et al., 2010; Jimenez et al., 2009; Chang et al., 2010; Duplissy et al., 2011).

2.1 Simulation chamber and associated instruments

CESAM is a 4.2 m³ cylindrical stainless steel chamber which has been designed to investigate both atmospheric gas-phase and aerosol-phase chemistry. As described by Wang et al. (2011), the wall properties and ventilation system guarantee a lifetime for long sub-micron particles and enables the study of aerosol aging for more than 20h. Water vapour can be directly injected in CESAM and thus the RH of the reaction mixture can be varied *in situ* from 0 to 100%.

The basic experimental setup and a schematic view of the CESAM chamber are shown in Figure 1. Temperature and relative humidity in the chamber are monitored using a Vaisala transmitter equipped with a capacitive thin-film Humicap sensor. The sensor was calibrated prior each experiments. The RH accuracy was $\pm 1.9\%$ up to 90% RH and the temperature accuracy was $\pm 0.1^\circ\text{C}$ at 20°C.

α -pinene, ozone and their reaction products were continuously monitored by a Fourier-Transform InfraRed spectrometer (FTIR) from Bruker GmbH (Ettlingen, Germany) coupled to a multi reflection cell allowing an optical path of 192 m. Additionally, ozone was monitored by a commercial Horiba APOA 370 instrument (Kyoto, Japan) with a detection limit of 0.5 ppb and a precision of 0.1 ppb.

The particle number size distribution was measured using a Scanning Mobility Particle Sizer (SMPS including a DMA 3080 and CPC 3010, TSI). The SMPS was operated at flow rates 3/0.3 Lpm (sheath flow/aerosol sample flow). The aerosol flow was diluted with filtered air before entering the CPC, in order to maintain the nominal flowrate at 1 Lpm in the CPC. The dilution air flow was sucked from the simulation chamber to avoid any pressure gradient in the SMPS. The SMPS scanning time was 2 minute 15 sec in total. The resulting measured size distribution ranged from 14 to 505 nm. The size calibration of the SMPS was performed using monodisperse

PolyStyrene Latex spheres (PSL, Duke Scientific). PSL, with diameters ranging from 100 to 500 nm, were nebulised with a constant output atomiser (TSI model 3076). The measured diameters were found to be larger than the PSL certified diameters (by about 10% for 100 nm PSL spheres), so a correction factor was applied. Size distributions were corrected by the SMPS software for both the loss by diffusion of particles in the SMPS tubing, the contribution of multicharged particles and the dilution of the aerosol flow before entering the CPC.

2.2 Optical properties measurements

The aerosol scattering coefficient (σ_{scat}) was measured using an integrating nephelometer (model M9003, Ecotech). The nephelometer operates at 525 nm wavelength and measures light scattered from particles at angles between 10° and 170°. It also measures temperature at both the sample inlet and within its cell with an accuracy of $\pm 0.6^\circ\text{C}$, and the RH within its cell with an accuracy of $\pm 3\%$. Prior to each series of experiments, the nephelometer was calibrated using particle-free air and CO_2 .

The particle light absorption coefficient (σ_{abs}) was determined by means of an aethalometer (Model AE31, Magee Scientific) operated with several light sources at seven wavelengths, covering the near ultra-violet to the near infrared wavelength range ($\lambda=370, 470, 520, 590, 660, 880$ and 950 nm). The aethalometer measures the attenuation of the transmitted light through a quartz fibre filter with increasing particle loading. This measurement can suffer from artefacts associated with reactions with oxidants occurring on particles deposited on the surface of the filter and desorption of gaseous compounds from the filter (Weintgartner et al., 2003). Thus, a charcoal denuder was installed upstream of the aethalometer to remove ozone and VOCs. It has also been observed that the aethalometer can suffer from biases at high RH as a result of the filter taking up water and scattering more light compared to the reference measurement (Cappa et al., 2008; Arnott et al., 2005). Therefore, the aethalometer was not used when the chamber RH was higher than 1%.

The spectral attenuation coefficient (σ_{attn}) resulting from the attenuation of light through the sampled aerosol on the filter was obtained using equation (1):

$$\sigma_{\text{attn}}(\lambda, m) = \frac{A}{Q} \frac{\Delta \text{attn}}{\Delta t} \quad (1)$$

where A is the spot area, Q is the volumetric flowrate and Δattn is the change in light attenuation during the time interval Δt . It is well known that σ_{attn} obtained with aethalometers are higher than true σ_{abs} (Arnott et al., 2005; Bond et al., 1999; Bond and Bergstrom, 2006; Cappa et al., 2008; Collaud Coen et al., 2010; Weingartner et al., 2003). Various systematic errors need to be corrected. The detailed description of the method used to correct σ_{attn} is shown in the Supplementary Material.

The details of the calculations of CRI of SOA under dry conditions ($\text{RH} < 1\%$) were given in Denjean et al. (2014a). The CRI was determined by optical closure experiments involving scattering and absorption coefficients (respectively σ_{scat} and σ_{abs}) measured at 525 nm and the number size distribution. Briefly, σ_{scat} and σ_{abs} were calculated according to equation (2):

$$\sigma_{\text{scat,abs}}(\lambda, m) = \int Q_{\text{scat,abs}}(D_p, \lambda, m) \frac{\pi}{4} D_p^2 \left(\frac{dN}{d \log D_p} \right) d \log D_p \quad (2)$$

where D_p is the geometrical particle diameter, $\frac{dN}{d \log D_p}$ is the number size distribution and $Q_{\text{scat,abs}}(D_p, \lambda, m)$ represent both the scattering and absorption efficiencies. $Q_{\text{scat,abs}}(D_p, \lambda, m)$ were calculated using Mie scattering calculations described by Bohren and Huffman (1983).

The values of σ_{scat} and σ_{abs} of SOA were simultaneously measured and compared with those calculated based on the Mie theory. To allow this comparison, all measured σ_{scat} were corrected from the sample temperature and pressure and from the angular truncation error examined for the nephelometer. The best-guess CRI was determined by minimizing the difference between measured σ_{scat} and σ_{abs} and those obtained using Mie calculations.

2.3 Hygroscopic properties measurements

The hygroscopic properties of SOA were investigated using two complementary approaches as described in details in Denjean et al. (2014a). Briefly, (1) a Hygroscopic Tandem Differential Mobility Analyzer (H-TDMA) was used to measure the RH dependency of the particles diameter for single size particles and (2) *in situ* experiments within the CESAM chamber allowed for the measurement of σ_{scat} changes after water uptake for polydisperse aerosol size distribution.

1

2 The hygroscopic behaviour of monodisperse SOA was measured with a H-TDMA. The
 3 instrument is composed of a DMA (TSI, 3080) that selects an initial mobility diameter under dry
 4 conditions $D_{p,m}(RH_{dry})$, an aerosol humidifier with a controlled higher RH and a DMA (TSI,
 5 3080) coupled to a CPC (TSI, 3010) to measure the wet size distribution over mobility diameter
 6 $D_{p,m}(RH)$. Both DMAs of the H-TDMA were calibrated using monodisperse PSL particles (Duke
 7 Scientific) ranging from 100 to 500 nm. A size shift was observed for both DMAs. The particles
 8 residence time for humidification is 15 seconds and corresponds to the residence time in the
 9 aerosol humidifier plus the transit time in the second DMA. In this study, the H-TDMA was
 10 typically operated at a constant high RH ($\pm 1\%$) of 90%. The hygroscopic size growth factor
 11 (GF) describes the relative increase in the geometric diameter of particles due to water uptake at
 12 a specific RH according to equation (3):

$$GF(RH) = \frac{D_{p,m}(RH)}{D_{p,m}(RH_{dry})} \quad (3)$$

13 Dry and humidified mobility diameters were obtained by assuming that the size distributions
 14 exhibited a lognormal profile. H-TDMA measurements were validated by measuring the
 15 humidogram of laboratory generated ammonium sulfate particles. The GF was found to agree
 16 with values calculated using the Köhler model (Denjean *et al.*, 2014a).

17

18 In addition to H-TDMA measurements, the hygroscopic behaviour of polydisperse SOA was
 19 measured by exposing the particles *in situ* to increasing humidity in the CESAM chamber. Water
 20 vapour produced from a small glass vessel filled with ultrapure water (18.2MΩ, ELGA Maxima)
 21 was injected into the chamber and mixed therein by the chamber's stainless steel fan. The RH in
 22 the chamber increased linearly from 0 to 100% within approximately 1 hour. The change in σ_{scat}
 23 due to hygroscopic growth of the overall particle size distribution was monitored by the
 24 nephelometer. The scattering growth factor ($f(RH)$) is the ratio between the scattering coefficient
 25 at a specific RH ($\sigma_{scat}(RH)$) and the dry scattering coefficient ($\sigma_{scat}(dry)$) (equation 4):

$$f(RH) = \frac{\sigma_{scat}(RH)}{\sigma_{scat}(dry)} \quad (4)$$

26 During humidification, RH was monitored within the chamber and at the inlet of the
 27 nephelometer. With this approach, the residence time for particle humidification was a few

minutes which is significantly longer than during H-TDMA measurements. It has been previously shown (Denjean et al., 2014a) that the two approaches of hygroscopicity measurements could lead to different results, giving information on water transfer dynamics, possible particles reorganisation or phase transfer equilibrium.

2.4 Aerosol sampling and chemical composition analysis

A High Resolution time-of-flight aerosol mass spectrometer (AMS, Aerodyne) (DeCarlo et al., 2006) was used to determine the bulk composition of organic aerosols (Aiken et al., 2007; Aiken et al., 2008). The instrument was used under standard conditions (vaporiser at 600°C and electron ionisation at 70 eV). Standard AMS calibration procedures using NH_4NO_3 particles were performed throughout the campaign and included Ionisation Efficiency (IE) calibration using Brute Force Single Particle (BFSP) method at 350 nm particle mobility diameter as well as size calibration using DMA-selected dried particle sizes over the range of interest 100-350 nm. Single ion calibration as well as baseline and threshold were analysed prior each experiment.

The AMS was operated during three experiments. The instrument was switched between two modalities: a single-reflectron configuration (V-mode) which offers a higher sensitivity and lower resolving power (up to ~2100 at m/z 200) and a double-reflectron configuration (W-mode) which provides a higher resolving power (up to ~4300 at m/z 200) but a lower sensitivity (De Carlo et al., 2006). Data were taken with a time resolution of 8 minutes. In V-mode, data were collected in the mass spectrum (MS) mode for 5 minutes, for aerosol quantification, and in the particle time of flight (PToF) mode for 1 minute for size distribution measurement (De Carlo et al., 2006, Canagaratna et al. 2007). Only MS data were recorded in the W-mode (2 minutes).

The AMS data were analysed using Squirrel (ToF-AMS Analysis 1.51B) and Pika (ToF-AMS HR Analysis 1.10B) packages for the software Igor Pro 6.21 (Wavemetrics, Inc., Portland, OR, USA). Default Collection Efficiencies (CE) and Relative Ionisation Efficiencies (RIE) were used for quantification of SOA. Air interferences were removed by adjusting the fragmentation table (Aiken et al., 2008; Allan et al., 2004). High-resolution analyses (De Carlo et al., 2006) were performed on V-mode data by integrating each $\text{C}_x\text{H}_y\text{O}_z$ ion in the mass range 12–180 m/z while W-mode data were used only to check for possible interferences. Elemental ratios (O:C and H:C)

1 were calculated according to the procedure described by Aiken et al. (2007; 2008). Measurement
2 uncertainties were estimated to be $\pm 30\%$, as determined by Aiken et al. (2008).

3
4 Further chemical analyses were performed by collecting SOA on PTFE filters (Zefluor, 47mm
5 diameter, 2 μm pore size, Pall Life Sciences), which were cut to the size of the collector using
6 ceramic scissors. A stainless-steel support was used for sampling and concentrating the particles
7 in a small filter area (0.9 mm^2). An active charcoal denuder was installed upstream of the filter
8 to remove ozone and VOCs. Before sampling, filters were extracted three times for 10 minutes
9 with dichloromethane (99.8%, HPLC grade) in an ultrasonic bath and baked for 5h at 250°C .
10 Filter samplings were performed in the chamber at different reaction times from 30 minutes to
11 17 hours at a nominal flow rate of 2 Lpm for a sampling time varying between 30 min to 2 hours,
12 depending on the total SOA volume concentration.

13
14 X-ray photoelectron spectrometry (XPS) was used to quantify the O:C of the SOA surface to a
15 depth less than 10 nm. Measurements were performed on a VG ESCALAB 250 instrument using
16 monochromatic Al K_α radiation (1486.6 eV). The X-Ray spot size was 500 μm . The analysis
17 chamber of the instrument was maintained at pressures ranging between 10^{-8} and 10^{-11} mbar.
18 Survey spectra of SOA were measured over an 1100 eV range at a resolution of 1 eV per step
19 and 100 eV pass energy (Figure S1a in the Supplementary Material). All peaks were referred to
20 the C_{1s} binding energy at 285.0 eV. The high-resolution C_{1s} spectrum showed the presence of C,
21 O and F. The deconvolution of the C_{1s} signal was performed in peaks corresponding to the $-\text{CO}_2$
22 (indicative of carboxylic acids, peroxides), C-O (alcohols, aldehydes, carboxylic acids,
23 peroxides, ethers), C-C/C-H (aliphatic functional group) and the C-F bonds of the PTFE filter in
24 order to optimise the fit (Figure S1b). Quantification of the O:C was performed using the
25 integrated peak areas of O_{1s} spectra, the peaks area of the $-\text{CO}_2$, C-O and C-C/C-H from the C_{1s}
26 spectra and the manufacturer's sensitivity factors.

27 28 **2.5 Experimental protocol**

29 Experiments were carried out without any seed particles or OH radical scavenger. Prior to each
30 experiment, the CESAM chamber was evacuated overnight to typically 4×10^{-4} mbar. The
31 chamber was then filled to atmospheric pressure with a mixture of 200 mbar of Oxygen (Air
32 Liquide, Alphagaz class 1, purity 99.9%) and 800 mbar of Nitrogen produced from the

evaporation of a pressurised liquid nitrogen tank (Messer, purity > 99.995%, H₂O < 5ppm). Table 1 shows the experimental conditions of this study. Ozone was generated by a Corona discharge in pure O₂ using a commercial dielectric ozone generator (MBT 802N, Messtechnik GmbH, Stahnsdorf, Germany). After ozone concentration stabilisation around 250 ppb in the chamber, approximately 200 ppb of α -pinene was introduced. This injection was performed by flushing into the chamber (in an O₂ flow) the entire volume of a glass bulb containing a known pressure in a glass bulb of a known volume of pure gaseous α -pinene (Aldrich, 98%). Particles were formed immediately after α -pinene injection. During the experiments, RH was lower than 1% and the temperature ranged between 14°C and 27°C (for different experiments).

The total pressure in the chamber was maintained constant by adding gaseous nitrogen to compensate for the continuous sampling by the instruments. As a consequence, the N₂:O₂ ratio increased with the reaction time. After 20 hours of reaction, the hygroscopicity of polydisperse particles was investigated by injecting water vapour in the chamber. The RH in the chamber increased linearly from 0 to 100% within approximately 1 hour. The nephelometer was used to follow *in situ* the variation in σ_{scat} as a function of RH.

3. Results

3.1. Reaction profile and aerosol yield

Typical evolution of α -pinene, ozone, particle number size distribution and mass concentrations are shown in Figure 2. α -pinene was essentially consumed after 4 hours of reaction. Since no OH scavenger was used, α -pinene was oxidised by O₃ and also by OH radicals produced by the reaction. A rapid formation of SOA was observed as the α -pinene was oxidised. The aerosol mass reached its maximum of $130 \pm 39 \mu\text{g m}^{-3}$ (when most α -pinene was consumed) and remained constant thereafter.

Table 1 summarises the results for the experiments conducted. The aerosol yield (*Y*) was given by equation 5:

$$Y = \frac{\Delta[\text{SOA}]}{\Delta[\alpha\text{pinene}]} \quad (5)$$

where $\Delta[\text{SOA}]$ is the SOA mass concentration produced for a given amount of α -pinene reacted, $\Delta[\alpha\text{pinene}]$, both corrected for dilution. Since α -pinene was introduced after ozone into the

chamber, and due to the very fast consumption of α -pinene, its initial concentration could not be accurately measured. For the yield calculations, we used the SOA mass concentration and α -pinene concentration after 5 minutes of reaction as initial concentrations. SOA mass concentration was calculated from the SMPS number size distribution corrected for dilution and the measured particle density, assuming homogeneous spherical particles. The SOA density of 1.2 g m^{-3} was estimated from the mobility mode and aerodynamic mode obtained from the SMPS and AMS measurements respectively, as described by DeCarlo *et al.* (2004) and Katrib *et al.* (2005). No significant change in the SOA density was found during the experiment. Furthermore, the retrieved value is in agreement with previous laboratory studies (Shilling *et al.*, 2008; Saathoff *et al.*, 2009). Table 1 shows that the SOA mass concentration of SOA varied from 44 to $139 \text{ } \mu\text{g m}^{-3}$, resulting in calculated yields values ranging between 0.07 and 0.21. The comparison between the steady-state aerosol yields determined in this study and the values reported in the literature is available in the supplementary material (Figure S2).

3.2. Complex refractive index

The real and imaginary parts of the CRI were retrieved to describe the scattering and absorbing characteristics of SOA.

The imaginary part of the CRI was estimated from the attenuation measurements obtained with the aethalometer. Since the attenuation coefficient depends on the quantity of particles accumulated on the filter, the CRI could not be calculated during the SOA formation, while the mass concentration increased rapidly until the α -pinene was completely consumed. **Erreur ! Source du renvoi introuvable.** S3 (Supplementary Material) shows that the absorption coefficient in the 370-950 nm wavelength range estimated after 14 hours of reaction is close to or below zero, indicating that the SOA do not absorb radiation in the visible to near-UV region. Therefore, the imaginary part of the CRI was set to zero. This result is consistent with previous studies on the absorbing properties of α -pinene- O_3 SOA. Cappa *et al.* (2008) observed no significant absorption of α -pinene- O_3 SOA at 532 nm with a photoacoustic spectrometer. Schnaiter *et al.* (2003) and Nakayama *et al.* (2010) measured no significant difference between scattering and extinction coefficient in the visible spectrum. Nakayama *et al.* (2012) showed that the imaginary part of the CRI values was negligible (< 0.003) in the 405-781 nm wavelengths.

1 Recently, Liu et al. (2013) reported also an imaginary part of the CRI below 10^{-4} in the 355-781
2 nm wavelengths range.

3
4 Figure 3a shows the real part of the retrieved CRI at 525 nm as a function of reaction time. Real
5 CRI were calculated when SOA exhibited a diameter high enough to allow for the detection of
6 the full size distribution by the SMPS. Based on the measurement uncertainties of $\pm 10\%$ for σ_{scatt} ,
7 we estimate an absolute error associated with the real CRI of ± 0.02 . The real CRI decreased from
8 1.43-1.60 at the beginning of the reaction to 1.32-1.38 after 14 hours of reaction. In contrast, no
9 significant changes in the CRI were observed during the first hour of reaction for experiments
10 E050210 and E080210. As shown in Figure 3b, the aerosol mass concentrations in these
11 experiments were stabilised after 20 minutes of reaction, which was significantly faster than in
12 the other experiments. This may result from differences in the $[\text{VOC}]/[\text{O}_3]$ concentrations (see
13 discussion in section 3.1 on the initial concentrations of α -pinene), leading to different reaction
14 kinetics (Chan et al., 2007). The real CRI appears to change significantly in most of the
15 experiments between 10 minutes and 1 hour of reaction and it is possible that these variations
16 were not detected in experiments E050210 and E080210. Values for the real part of the CRI
17 retrieved in this study at 525 nm are compared with the literature values obtained for SOA
18 produced during α -pinene + O_3 SOA formation (Table 2). Because our experiments were
19 performed under various time scales and under various conditions and because our CRI values
20 evolved with SOA aging, we can globally say that our values (1.32-1.60) were in the same range
21 as the previous ones.

22 23 **3.3. Hygroscopic properties**

24 The first insight in the hygroscopic behaviour of SOA was brought by measuring humidograms
25 of SOA with the H-TDMA. Particles of 115 nm diameter were selected for the analysis after 2
26 hours aging and 190 nm after 14 hours aging. The GF uncertainty was estimated from the
27 uncertainty in retrieving the geometric diameters and the RH uncertainty was based on the
28 weighted average of the RH sensors uncertainties at the entrance of the H-TDMA. The GF values
29 at 90% RH of SOA are reported in Figure 4 as a function of time. GF ranged from 1.02 (± 0.02)
30 to 1.07 (± 0.02) at $90 \pm 4.2\%$ RH during and after SOA formation, in agreement with values
31 previously reported in the literature (Table 3).

The comparison of the mobility diameter as a function of RH after 2 and 14 hours of SOA aging is shown in figure 5a. Unusual behaviour of the mobility diameter and $f(RH)$ at the end of the experiment (at $t=16-20h$) is evident. Mobility diameters dropped to a minimum between 20-50% RH for SOA after 14 hours reaction that was not observed for “fresh” SOA. Figure 5b shows the RH dependence of $f(RH)$ of the polydisperse size distribution as measured directly in the chamber during water injection after 1 hour and 16 hours of the reaction. Two RH scales are used: one for the measurements performed within the chamber and the other for the measurements performed using the nephelometer. In fact, a small drying of 10% has been observed between the chamber and the nephelometer. No experiment with nephelometer measurements at the beginning and the end of the same date was available. However, the experiments were performed under very similar initial conditions as shown by the reproducible values of CRI and GF that are also comparable with each other. There is a first increase of $f(RH)$ as RH increases from 0 to 30%.. This can be due to a change in the physical state of SOA. In fact, particles with irregular shapes or porous should be observed at higher mobility diameters in the DMA than spherical and compact particles of the same mass (De Carlo et al., 2004). After water uptake, the particles may become more spherical or compact, leading to a shift of the mobile size distribution measured with the DMA (Milkhailov et al., 2009). Particles are expected to exhibit also a higher mass concentration, and the σ_{scatt} should increase after humidification (Adachi et al., 2011). This hypothesis is discussed in details in section 4.2. A second slope is observed between 80 and 90% RH, linked to the sharp increase of GF due to particle water absorption. All the experiments exhibited the same trend, but different values are observed at 90% RH. This can be explained by the different size distribution from one experiment to another. In particular, the proportion of particles larger than 100 nm is different resulting in a different capacity in absorbing water (Biskos *et al.*, 2006) and hence varying the observed $f(RH)$.

An important concern in measuring the hygroscopic properties of the particles is to allow sufficient time for particle-water vapor equilibrium. Various studies have discussed the possibility that insufficient time for humidification could result in an underestimation of the particles’ water content (Chan and Chan, 2005; Saxena et al., 1995; Duplissy et al., 2009; Denjean et al., 2014a). In this study, the residence time for SOA particles in the wet air stream was significantly longer when particles were humidified in-situ in the chamber. It took ~1 hour in the chamber, instead of 15 seconds in the H-TDMA. These two approaches can thus be

complementary to carry information on water transfer dynamics of α -pinene- O_3 SOA. We used Mie scattering calculations for homogeneous spheres to determine GF from $f(RH)$. σ_{scat} was calculated for different GF at specific RH. The optimal GF as a function of RH was determined so that the difference between measured σ_{scat} and those obtained using Mie calculations were minimized. Particles were assumed to be homogeneous spheres of uniform CRI. The CRI calculations were based on volume weighted refractive indices of SOA and water. Uncertainties on the theoretical GF were estimated from the standard deviation of the measured $f(RH)$, the uncertainties on the $f(RH)$ measurements and the uncertainties on the RH measurements. Figure 6 shows the comparison between measured and predicted GF values for SOA at two different reaction times: for “fresh” SOA (after 1 hour of reaction), and for “aged” SOA (after 16 hours of reaction). For both reaction times, the model approach agrees well with the measurements above 30% RH. It indicates no kinetic limitations of α -pinene- O_3 SOA for water uptake. For “aged” SOA (after 16 hours of reaction), the underestimation of the model below 30 % RH (Figure 6) might be due to a change in the physical state of SOA, as discussed in detail in section 4.2.

3.4. Chemical composition

The O:C of the bulk aerosol measured online by the AMS and the O:C at the surface obtained by XPS analysis on filter samples are shown in Figure 7. The bulk O:C ranged between 0.68 (± 0.20) to 0.55 (± 0.16). The uncertainties in bulk O:C given by Aiken et al. (2007) may be overestimated compared to the experimental variability and even experimental reproducibility observed in this study. In fact, we estimated the experimental uncertainties to be ± 0.01 from the standard deviation of the experimental values.

Two potential processes affecting the bulk O:C during the SOA formation are the gas/particle phase partitioning and chemical reactions in the condensed phase. The evolution of the AMS mass spectra during the reaction is shown in Figure S4 (Supplementary Material). The AMS mass spectra were dominated by m/z 44 for ‘fresh’ SOA, while the strongest signal was observed at m/z 29 for ‘aged’ SOA. Figure S5 also shows that f_{44} decreased while f_{43} increased with time. These observations indicate an increase of less oxidized semi-volatile compounds in the bulk particle phase with aging. This trend can be explained by an increasing partitioning of less oxidised semi-volatile compounds as the aerosol grows according to the structure-activity correlation between the oxygenated functional groups and the species vapour pressure (Pankow

1 and Asher, 2008). Additionally, some oligomerisation/dehydration, such as aldol condensation
2 or ester formation, can alter the SOA O:C and f_{44} (Camredon et al., 2010; Ziemann and Atkinson,
3 2012). Additional reactions in the particle phase can take place after the SOA formation, and can
4 for example lead to the formation of oligomers (Kalberer et al., 2004; Gao et al., 2004b; Tolocka
5 et al., 2004; Gao et al., 2004a; Yasmeen et al., 2012; Reinhardt et al., 2007; Rudich et al., 2007),
6 but do not appear to be associated with a significant change in the SOA functionality.

7
8 Bulk O:C obtained in the literature from different techniques are shown in Table 4. The bulk O:C
9 retrieved in our study are in agreement with Shilling et al. (2008), Chhabra et al. (2010),
10 Reinhardt et al. (2007) and Tolocka et al. (2006). Values at the beginning of the reaction are,
11 however, higher than those obtained by Aiken et al. (2007) and Qi et al. (2012) who used AMS
12 to retrieve the O:C. Limited information is given by Aiken et al. (2007) about their experimental
13 conditions. In particular, temperature and reaction time scales are critical parameters which can
14 significantly affect the chemical composition of SOA (Warren et al., 2009; Shilling et al., 2008;
15 Chhabra et al., 2010). Qi et al. (2012) did not provide any uncertainty in their AMS
16 measurements, but taking into account the uncertainties of $\pm 30\%$ given by Aiken et al. (2007),
17 their O:Cs are in agreement with our values.

18
19 The O:C at the surface of the SOA is significantly lower than the one in the bulk (Figure 7) and
20 ranges between 0.33 (± 0.07) and 0.46 (± 0.08). Some volatile compounds could evaporate during
21 the XPS analysis due to the low pressure in the instrument. In our companion paper (Denjean et
22 al., 2014b), an increase of the O:C has been observed after heating by the evaporation of semi-
23 volatile components of SOA. Therefore, the O:C at the surface of the SOA was certainly
24 overestimated in the present study. This suggests that the SOA composition is not homogeneous,
25 but composed of less oxidised species at its surface. To our knowledge, our study is the first one
26 to investigate the surface O:C of the SOA and thus no comparison with literature is possible.
27 Nevertheless, this observation is in very good agreement with one of the hypotheses by McIntire
28 et al. (2010). The authors proposed a hydrophobic shell model for hydrocarbon + O_3 SOA due to
29 the burying of R-COOH compounds and other polar groups inside the particle. Here, we report
30 direct experimental measurements which support this hypothesis.

4. Discussion

4.1. Dependence of the CRI and GF on the chemical composition

Our results suggest that the real part of the CRI is closely related to SOA chemical composition. Figure 8 shows that the real part of the CRI decreases substantially as the O:C of the bulk SOA decreases. This trend is in agreement with the positive correlation between the extinction coefficient at 532 nm wavelength and O:C observed by Cappa et al. (2011) for the heterogeneous OH oxidation of squalane and azelaic acid particles. Nakayama et al. (2013) reported increasing real part of the CRI at 532 nm with increasing O:C for SOA generated from the photooxidation of toluene. Recently, Flores et al. (2014) showed that the real part of the CRI at 405 nm is positively correlated to the O:C for a mixture of biogenic VOC (α -pinene and limonene) and biogenic VOC mixture with subsequent addition of an anthropogenic VOC (p-xylene-d₁₀). On the other hand, an opposite trend was observed by Nakayama et al. (2012) with a decrease of the real part of the CRI at 532 nm wavelength with increasing O:C for SOA produced from the ozonolysis and photooxidation of α -pinene. Lambe et al. (2013) studied SOA formed by OH-oxidation of various gas-phase precursors (naphthalene, tricyclo[5.2.1.0^{2,6}]decane, guaiacol and α -pinene) and retrieved a decreasing real part of the CRI with increasing oxidation. These different trends suggest that the change of the real CRI as a function of O:C depends strongly on the organic aerosol source and aging in the atmosphere.

The decrease of bulk O:C is expected to induce a decrease in SOA hygroscopic properties (Massoli et al., 2010; Jimenez et al., 2009; Chang et al., 2010; Duplissy et al., 2011). However, the GF remained surprisingly constant over 20 hours of reaction (Figure 4). This trend contradicts with Saathoff et al. (2003) and Cocker et al. (2001) who reported an increase of the GF from 1.080 to 1.106 and 1.065 to 1.1 respectively within time scale of 6 hours. However, our constant GF is consistent with the constant hygroscopicity observed by Warren et al. (2009) and Qi (2010) for α -pinene-O₃ SOA. The constant GF observed in our study may result from: 1) the sensitivity of the instruments. The GF was studied after 2 hours of reaction when the decrease of the oxidation degree was slow. The O:C of bulk SOA decreased from 0.65 to 0.60 between 2 and 10 hours aging (Figure 7), that should result in a GF decrease of 0.03 according to the linear GF-to-O:C relationship reported by Massoli et al. (2010). This expected GF variation (due to the

chemical composition change) is smaller than the sensitivity of the H-TDMA. 2) Other factors than the bulk O:C can control the water uptake of SOA: recently, Alfarrar et al. (2014) reported a positive correlation between hygroscopicity of particles and their degree of oxidation for SOA produced from the photooxidation of α -pinene, β -caryophyllene, linalool and myrcene, but not for limonene SOA. They suggested that other factors such as solubility, surface tension, molecular weight, density and particle phase are likely to be playing important roles in controlling GF values. We observed that the O:C of the particle surface was lower than the bulk O:C and remained constant over time (see section 3.4.1.). This structure would affect the heterogeneous chemistry of the particle and avoid the water absorption on particles. Broekhuizen et al. (2004) observed that the oxidation of oleic acid by low concentration of ozone, shown to produce oxygenated products (Moise and Rudich, 2002), did not increase the particle CCN activity. Katrib et al. (2004) demonstrated that hydrophobic products were formed at the surface of oleic acid particles after ozone exposure that prevented from water adsorption on particles. McIntire et al. (2010) exposed particles formed from ozonolysis of surface-bound alkenes to ozone and observed formation of polar groups buried inside a hydrophobic shell. This is also consistent with Moussa et al. (2009) who observed that the uptake of water does not increase after the oxidation of surface-bound alkenes. A core-shell structure could explain the low O:C of the surface α -pinene-O₃ SOA obtained in the present study and could affect significantly the heterogeneous chemistry of the particle.

4.2. Effect of RH and aging on the viscosity of SOA

An unexpected effect of aging on the dynamic of the size distribution evolution during humidification has been observed (section 3.3 and Figure 5). For the 16 hours old SOA, humidification led to a decay of the mobility diameter and an increase of σ_{scatt} while such behaviour was not observed for the 1 hour old of reaction. Although evaporation of semi-volatiles in the chamber could lead to a decrease of the observed mobility diameter, it cannot explain the increase of $f(\text{RH})$ below 40% RH. Furthermore, the observed decrease of the mobility diameter cannot be attributed to losses of particles to the walls of the chamber, since it would lead to a decrease of $f(\text{RH})$.

As it is difficult to believe that a RH increase would induce a sudden loss of matter for the particulate phase, it is hypothesised that the particles experienced a rapid change of shape when

humidified. Being observed for aged SOA only, these behaviours support a re-shaping of coagulated particles and an increase in sphericity for $RH > 20-30\%$. This implies that, below these RH values, coagulated particles were not spherical and hence they were not liquid enough for coalescence. Electron microscopy analysis was used to investigate the shape of SOA particles. SOA particles after 1 hour and 14 hours of reaction were analysed by transmission electron microscopy (TEM) (Figure S6 in the Supplementary Material). Experimental details are given in the Supplementary Material. In total 50 particles were analysed by TEM. Figure S6 shows an example of SOA particles after 1 and 14 hours of reaction. Only spherical particles have been observed in the samples. This indicates that the coagulation taking place at the beginning and after 14 hours of reaction resulted in spherical coalesced particles.

A few recent studies have suggested that α -pinene- O_3 SOA is likely to be in an amorphous semi-solid or in an amorphous solid (glassy) state under dry conditions ($RH < 30\%$). Virtanen et al. (2011) reported that α -pinene- O_3 SOA bounces off impactor plates as if particles were glassy. Other studies reported that the evaporation kinetics of α -pinene- O_3 SOA was lower than expected by models for liquid droplets, indicating a nonliquid-like state (Cappa and Wilson, 2011, Vaden et al., 2011). These findings are in agreement with Perraud et al. (2012) who has shown that the uptake of organic nitrates by α -pinene- O_3 SOA did not follow absorptive equilibrium partitioning theory, as indicative of non-liquid-like behaviour. Moreover, some recent studies reported viscosity changes of α -pinene- O_3 SOA with RH . Saukko et al. (2012) observed lower bounce behaviour after the SOA was exposed to $RH > 50\%$, indicating a liquid-like, less viscous phase. Renbaum-Wolff et al. (2013) reported that the viscosities of α -pinene- O_3 SOA correspond to a semisolid or solid for $RH \leq 30\%$, a semisolid for $40\% \leq RH < 80\%$ and a liquid for $RH \geq 80\%$. This is consistent with our observation of the apparent step in mobility diameter and σ_{scatt} changes observed between 0 and 30% RH for SOA after 16 hours of reaction (Figure 5), which may be due to phase transition from a predominantly glassy state to a predominantly liquid state.

In addition, we find no apparent transition step between 0 and 30% RH for SOA after 1 hour of reaction. This observation mirrors the effect of aging on the physical state of SOA. We discussed previously the possible formation of oligomers in the particle phase during the formation of SOA. Roth et al. (2005) suggested that ambient organic particles could be in an extremely high viscosity, glassy state due to the presence of oligomeric constituents. There is evidence from

different organic materials that oligomerisation may lead to an increase in viscosity and potentially to the formation of glassy states (Koop et al., 2011). Abramson et al. (2013) reported decreasing evaporation kinetics during the aging of α -pinene- O_3 SOA, indicating that hardening occurs with time. Our observations are consistent with an increase in the SOA oligomer content with aging, leading to the transformation of SOA into glassy states under dry conditions.

These findings are of major importance as it suggests that particles may undergo reorganisation after the condensation of oxidised species. It is also supported by the core-shell structure of SOA observed with less oxidised species at its surface (section 3.4.1. and Figure 7). Due to possible kinetic limitations, the particles viscosity can strongly affect a number of key physical and chemical properties, such as water uptake, equilibrium partitioning between the gas and the particle phase or heterogeneous chemistry (Koop et al., 2011; Zobrist et al., 2008, Milkhaïlov et al., 2009, Murray 2008; Shiraiwa et al., 2013) and thus has considerable implications on the understanding of the impact of SOA on climate.

5. Implication for direct radiative effect and conclusions

The imaginary part of the CRI values for α -pinene- O_3 SOA was found to be negligible in the visible spectrum, indicating that these particles have a pure scattering effect. The single scattering albedo w_0 , a key parameter to estimate the influence of aerosols on the radiative balance, is estimated to be equal to 1, indicating that these particles have a climate cooling effect. Therefore, α -pinene- O_3 SOA do not have a direct aerosol heating potential.

In order to estimate the evolution of the direct radiative effect of SOA, the mass extinction efficiency k_{ext} (dry) at 525 nm of dry SOA calculated as the ratio of the measured σ_{scat} to the SOA mass concentration is shown in Figure 9. The k_{ext} values depend on the competitive effect of aerosol size and chemical composition. The decrease of the real CRI resulting from the changing particle chemical composition is expected to decrease k_{ext} while the size increase of the SOA from the nucleation mode to the accumulation mode should increase k_{ext} . At the beginning of the SOA formation, we obtained k_{ext} (dry) = $0.64 \pm 0.27 \text{ m}^2 \text{ g}^{-1}$, that is significantly lower than obtained after 14 hours of reaction where k_{ext} (dry) = $1.68 \pm 0.50 \text{ m}^2 \text{ g}^{-1}$. The mass extinction efficiency k_{ext} is also strongly affected by the ambient relative humidity (Figure 9). α -pinene- O_3

SOA did not show deliquescence below 90% RH, but a continuous water uptake with increasing RH. A constant GF (90% RH) was obtained throughout the experiment. From the measured $f(90\% \text{ RH})$ and $GF(90\% \text{ RH})$ obtained in this study, we can estimate σ_{scat} and SOA mass concentration at 90% RH and then $k_{\text{ext}}(90\% \text{ RH}) = 2.36 \pm 0.70 \text{ m}^2 \text{ g}^{-1}$. Water uptake by SOA at 90% RH produces an increase of k_{ext} by 40%. This simple calculation highlights the importance of taking into account the actual ambient conditions that the α -pinene- O_3 SOA experiences during its lifetime when estimating its induced direct radiative effect in the atmosphere.

For comparison, k_{ext} of organic aerosols calculated using the CRI and GF as prescribed in global models are shown in Figure 9. Mie scattering calculations for homogeneous spheres were performed to calculate k_{ext} . Number size distribution representing biogenic aerosol particles in the Amazon Basin (Martin et al., 2010) was used as an input in the Mie scattering calculations. For wet conditions, the CRI calculations were based on volume-weighted CRI of values of α -pinene- O_3 SOA and water. k_{ext} of SOA generated from a mixture of biogenics (mostly terpenes) released directly from plants (Lang-Yona et al, 2010) is also shown in Figure 9. Despite the wide range of variability of the CRI values assumed in global models for organic aerosols (1.45-0i to 1.6-0.003i), the values for k_{ext} were very similar for all models. The k_{ext} values retrieved in the present study for SOA after 14 hours of reaction agree with the values used for global models, while k_{ext} obtained for SOA at the beginning of the reaction are 2-7 times smaller. This work is a first attempt to assess the evolution of k_{ext} of SOA during its formation and aging. Our results suggest that the k_{ext} of atmospheric SOA is not static and a single size distribution and CRI does not appear sufficient to accurately model its direct radiative effect.

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11

Table 1. Initial Conditions, Temperatures, Relative Humidities, and Results of the Experiments.
All experiments started with ~ 200 ppb of α -pinene.

Experiments	O ₃ ^a (ppb)	Temp (°C)	RH ^b (%)	Mass concentration ^c ($\mu\text{g m}^{-3}$)	Yield ^d
E140910	251	23-25	<1	130	0.19
E050210	202	19-23	<1	117	0.18
E080210	232	19-20	<1	116	0.18
E090210	254	16-17	<1	118	0.15
E230910	246	22-23	<1	100	0.17
E160411	232	19-23	<1	70	0.16
E260411	247	20-25	<1	-	-
E300911	198	20-27	<1	78	0.11
E301111	197	18-20	<1	52	0.07
E091211	196	17-20	<1	81	0.11
E180411	201	21-26	<1	86	0.13
E280411	269	17-20	<1	-	-
E280911	293	22-27	<1	91	0.13
E021211	196	17-20	<1	70	0.14
E071211	162	16-19	<1	55	0.14
E200411	220	21-26	<1	97	0.15
E260911	292	23-26	<1	87	0.15
E281111	215	14-19	<1	44	0.13
E051211	185	16-19	<1	66	0.13
E120312	229	18-23	<1	139	0.21
E030512	216	20-23	<1	109	0.15
E060512	246	19-21	<1	102	0.16

^aOzone concentrations determined from FTIR measurement

^bRH before the injection of water at the end of the experiment

^cAerosol mass concentration estimated from the aerosol volume concentration corrected from dilution and by assuming a density of 1.2 g cm^{-3}

^dYield calculated from measured aerosol mass and the corresponding concentration of reacted α -pinene, both corrected from dilution. We used the SOA mass concentration and α -pinene concentration after 5 minutes of reaction as initial concentrations since the initial concentration of α -pinene could not be directly measured (see text).

Table 2. Comparison of real part of the complex refractive index (CRI) of α -pinene-O₃ SOA with previous studies.

Reference	Real CRI	Residence time	λ (nm)	$[\alpha\text{-pinene}]_{\text{initial}}$ (ppm)	$[\text{O}_3]_{\text{initial}}$ (ppm)
This study	1.60 (\pm 0.02) 1.33 (\pm 0.02)	10 mn 19 h	525	0.20	0.25
Kim <i>et al.</i> (2010)	1.45 (\pm 0.05)	2.5 h	670	0.50-5.00	0.10-1.00
Kim and Paulson (2013)	1.39 (\pm 0.02) 1.52 (\pm 0.02)	< 30 mn 4 h	532	0.13-0.17	0.50
Liu <i>et al.</i> (2013)	1.498 (\pm 0.002)	38 s	550	4.00	52.2
Nakayama <i>et al.</i> (2010)	1.41 (\pm 0.02)	2-3 h	532	0.10	2.00
Nakayama <i>et al.</i> (2012)	1.47–1.48 (\pm 0.02)	2-3 h	532	0.10	1.09-2.57
Redmond and Thompson (2011)	1.49 (\pm 0.04)	-	532	-	0.5-1.00
Schnaiter <i>et al.</i> (2003)	1.44	1.23 h	>350	0.06	0.47
Wex <i>et al.</i> (2009)	1.45	2 mn	visible	-	excess

1 Table 3. Comparison of size growth factor GF at 90% RH of α -pinene-O₃ SOA with previous
2 studies.

3

Reference	GF(90%RH)	Residence time	[α -pinene] _{initial} (ppm)	[O ₃] _{initial} (ppm)
This study	1.02 - 1.07 (± 0.02)	2 h	0.20	0.25
	1.02 - 1.07 (± 0.02)	20 h		
Prenni <i>et al.</i> (2007)	1.01 - 1.07 (± 0.02)	2 h	-	excess
Qi <i>et al.</i> (2010)	1.09	30 mn	0.05-0.10	0.30-0.34
	1.09	7 h		
Saathoff <i>et al.</i> (2003)	1.08 (± 0.01)	1 h	0.06	0.5
	1.11 (± 0.01)	6 h		
Warren <i>et al.</i> (2009)	1.02 - 1.16 (± 0.02)	30 mn	0.05	0.3-0.5
	1.02 - 1.16 (± 0.02)	6 h		

4

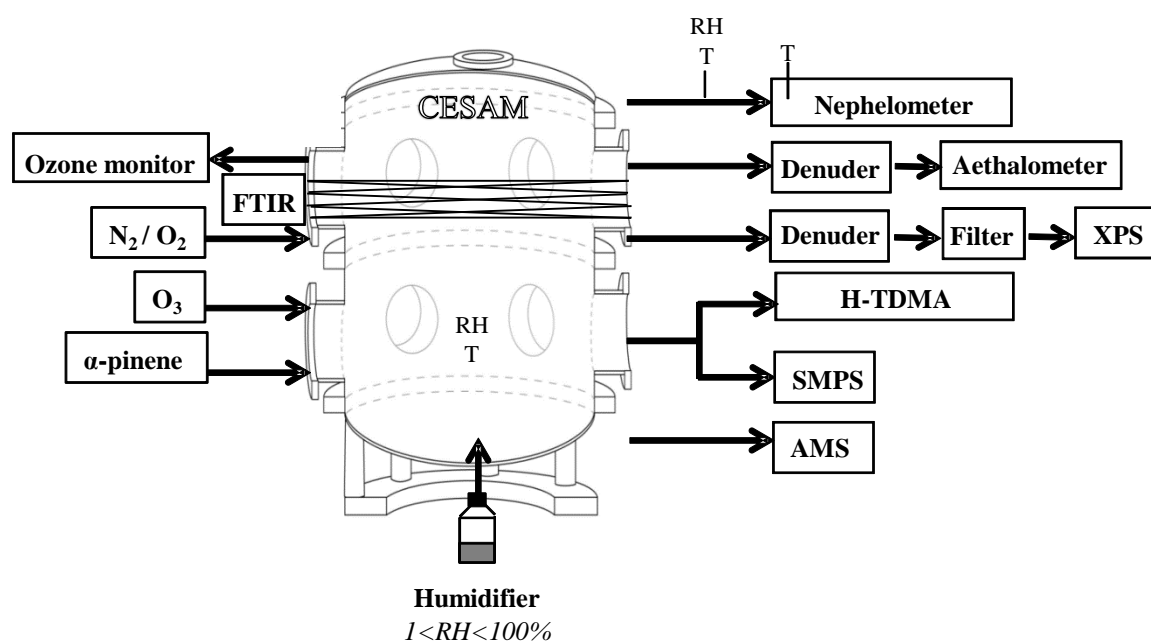
1 Table 4. Comparison of bulk O:C of α -pinene-O₃ SOA with previous studies in the literature.

2

Reference	Bulk O:C of SOA	Instrumentation
This study	0.55 (± 0.16) - 0.68 (± 0.20)	HR-ToF-AMS
Aiken <i>et al.</i> (2007)	0.27 (± 0.08)	HR-ToF-AMS
Shilling <i>et al.</i> (2008)	0.29 (± 0.09) - 0.45 (± 0.14)	HR-ToF-AMS
Chhabra <i>et al.</i> (2010)	0.43 (± 0.13)	HR-ToF-AMS
Qi <i>et al.</i> (2012)	0.33	HR-ToF-AMS
Reinhardt <i>et al.</i> (2007)	0.4 - 0.6	FTICR-MS
Tolocka <i>et al.</i> (2006)	0.37 (± 0.05) - 0.40 (± 0.12)	NanoAMS

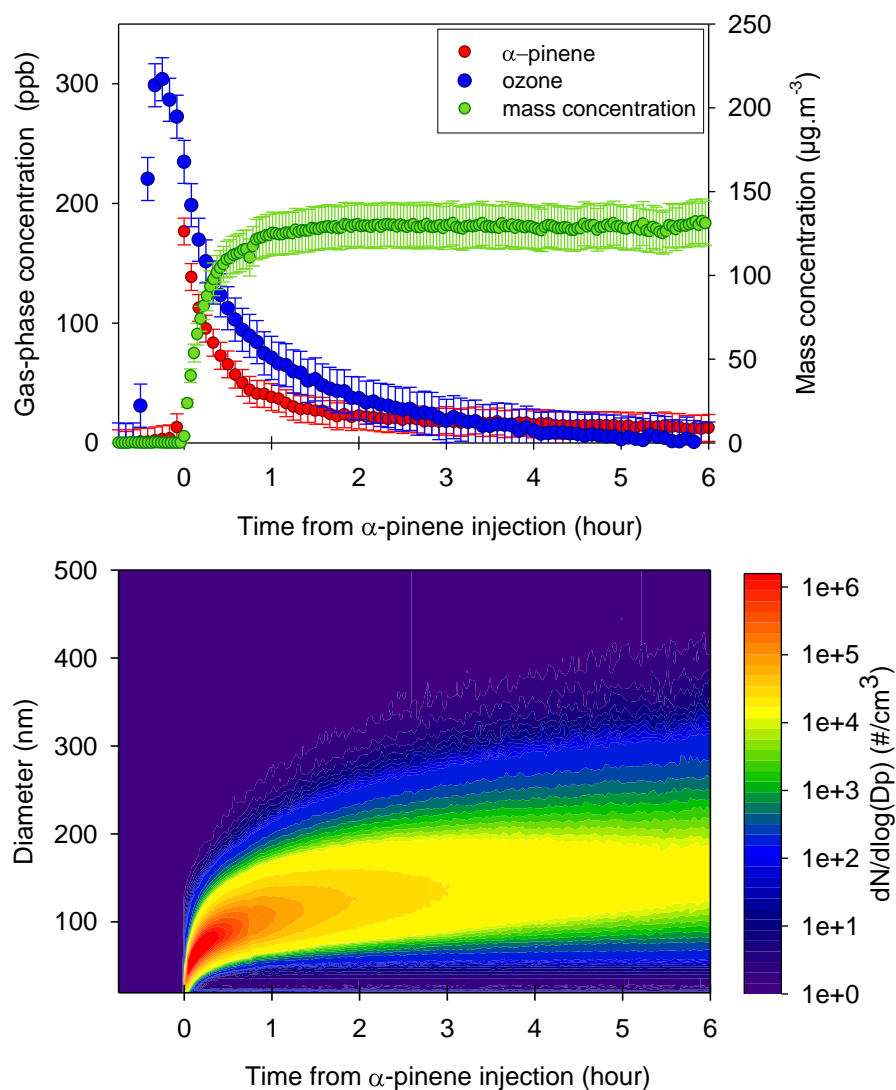
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- 1 Figure 1: Experimental set-up of the CESAM chamber used to measure aerosol chemical,
- 2 hygroscopic and optical properties.

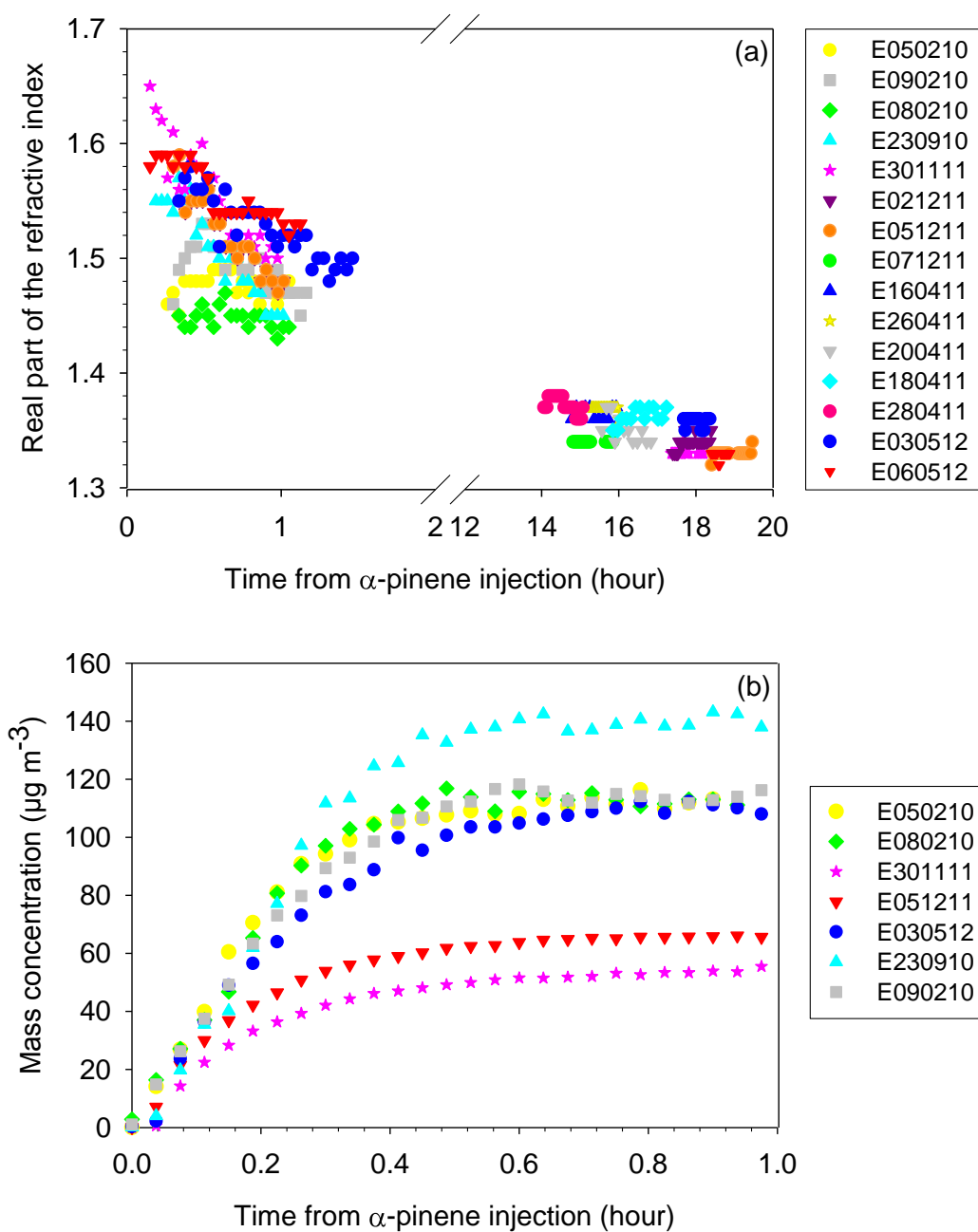


3

1 Figure 2. Temporal profiles of the initial phases of a typical α -pinene ozonolysis experiment
2 (E160410): (a) α -pinene (red), ozone (blue) and mass concentration of the SOA (green) and (b)
3 measured number size distribution.

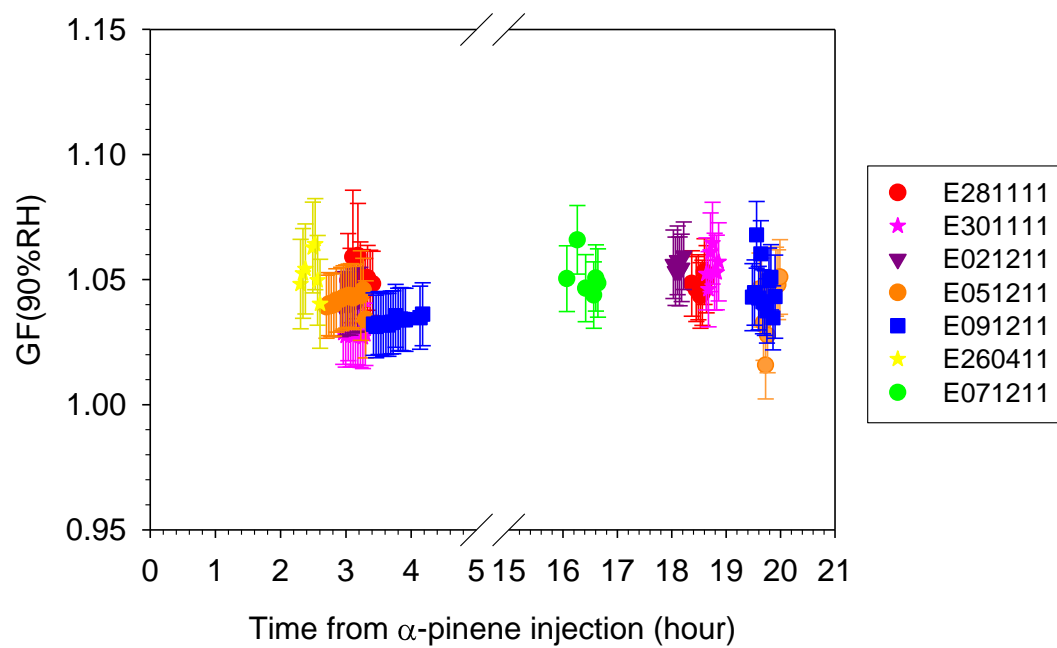


1 Figure 3: Calculated real part of the complex refractive index at $\lambda=525$ nm of SOA as a
2 function of the reaction time (upper panel). The lower panel shows the corresponding mass
3 concentration for the initial formation phase after injection (estimated particle density set to 1.2
4 g m^{-3})



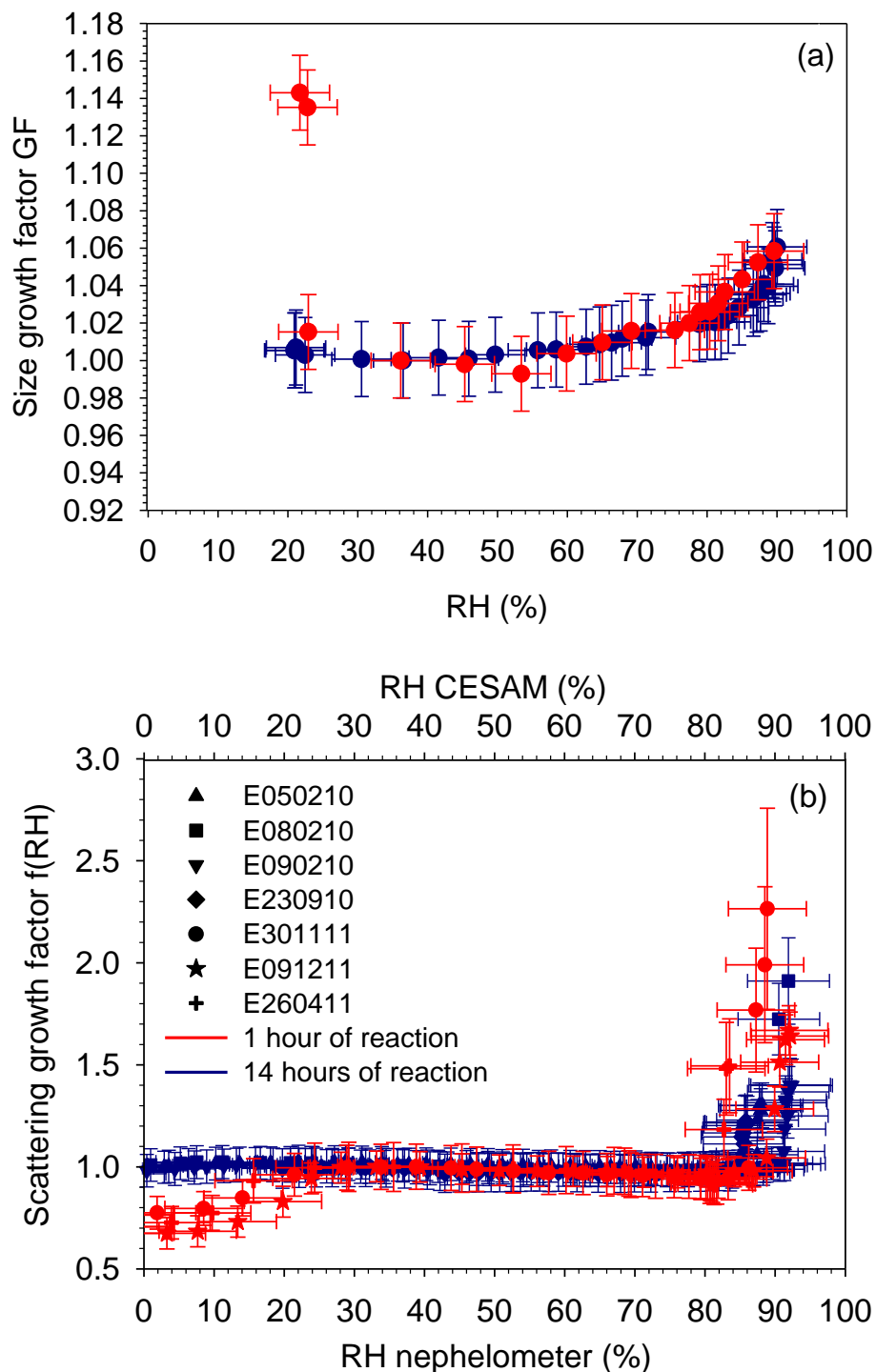
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1 Figure 4: Size growth factor GF at 90% RH obtained from the H-TDMA measurement as a
2 function of the reaction time.

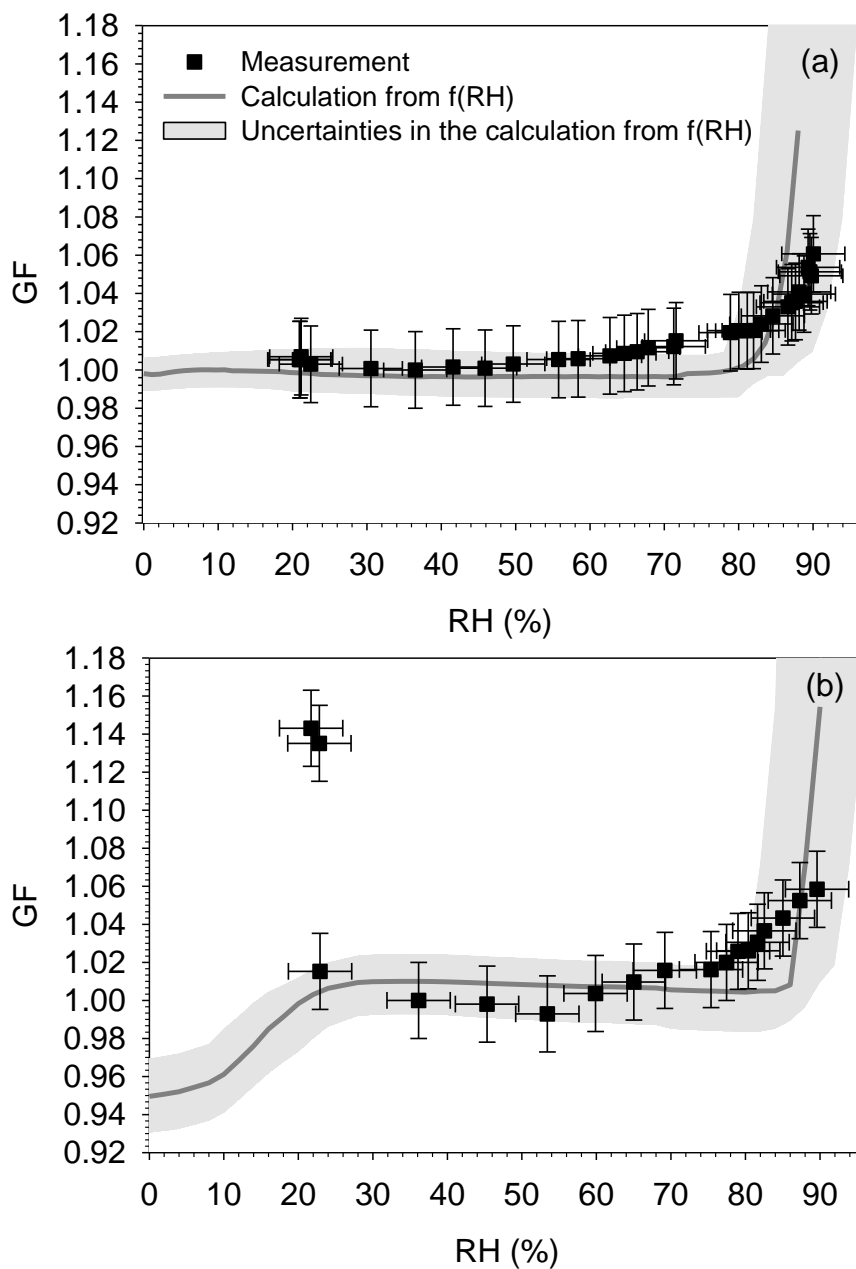


3

Figure 5: Humidograms of (a) size growth factor measured by the H-TDMA and (b) scattering growth factor measured by the nephelometer as a function of RH within the nephelometer (bottom axis) and the RH within the chamber (upper axis), for “fresh” SOA (after 1 hour of reaction) in red and for “aged” SOA (after 14 hours of reaction) in blue. For the calculations of $f(RH)$ and GF, we used $\sigma_{\text{scat}}(RH_{\text{dry}})$ and $D_{p,m}(RH_{\text{dry}})$ at 40 %.

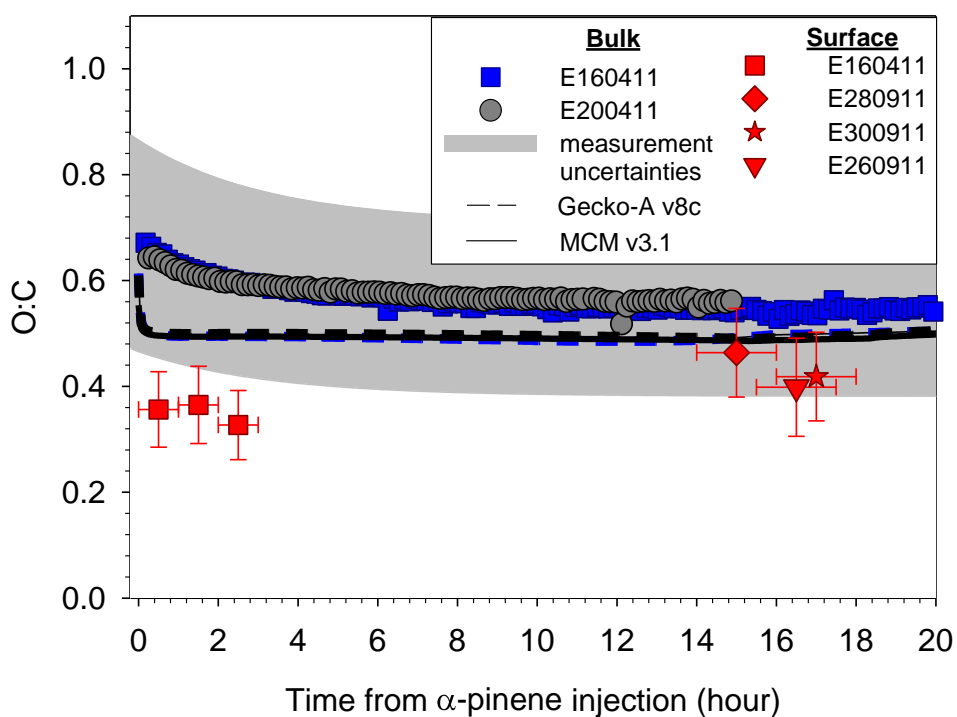


1 Figure 6: Humidograms showing measured (black symbols) and predicted GF (grey line) as a
2 function of RH of SOA (a) after 1 hour of reaction and (b) after 14 hours of reaction. The grey
3 area represent the uncertainties in the calculation of GF from $f(\text{RH})$.
4



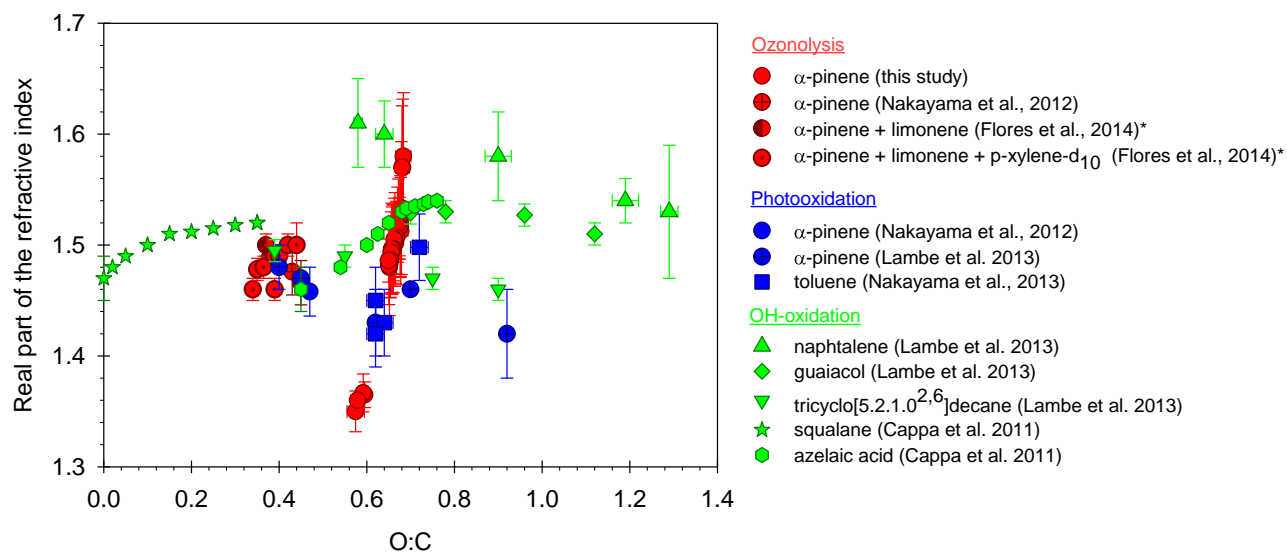
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1 Figure 7. O:C of the bulk SOA obtained from AMS measurements (blue and grey symbols) and
2 uncertainties in the bulk O:C values (grey area) as a function of time. These bulk O:C are
3 compared with those of the SOA surface determined from XPS analysis (red symbols).



4

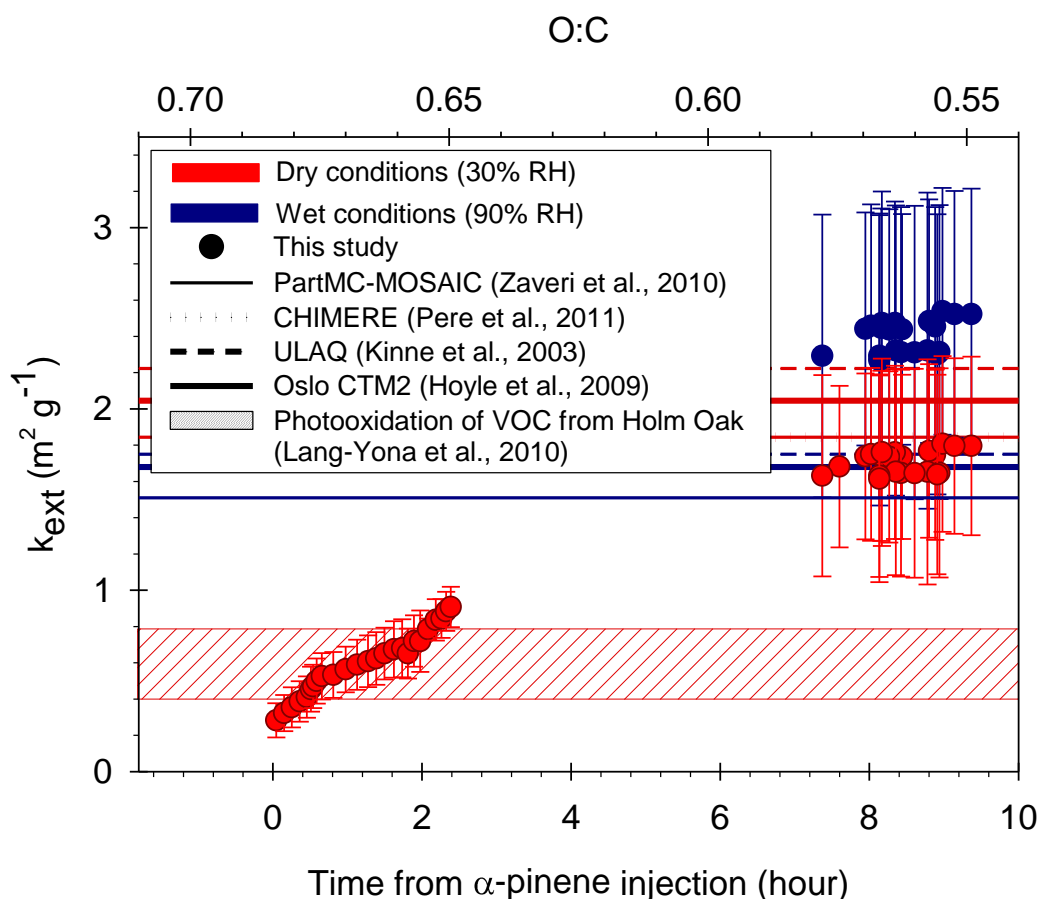
1 Figure 8: Real part of the refractive indices at $\lambda=525$ nm of SOA as a function of the bulk O:C.
 2 In our study, representative error bars represent $\pm 1\sigma$ in replicate measurement.



3
 4 * measured at $\lambda=405$ nm

5

1 Figure 9. Mass extinction efficiency (k_{ext}) at $\lambda=525$ nm of SOA under dry (red dots) and wet
 2 conditions (blue dots) as a function of time and bulk O:C. These k_{ext} values are compared to those
 3 used for organic aerosols in several global models (lines) and those of SOA generated from a
 4 mixture of biogenics released directly from plants (dashed area) . Zaveri et al. (2010) assumed a
 5 CRI of 1.45-0i and GF of 1.23 , Pere et al. (2011) a CRI of 1.45-0.001i, Kinne et al. (2003) a CRI
 6 of 1.60-0.003i and GF of 1.09 (for ULAQ model) and Hoyle et al. (2009) a CRI of 1.53-0i and
 7 GF of 1.03. The hygroscopic growth of SOA was not taken into account by Pere et al. (2011).



8

9