

# **Aqueous-phase photochemical oxidation and direct photolysis of vanillin—a model compound of methoxy-phenols from biomass burning**

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Manuscript submitted to *Atmospheric Chemistry and Physics*, 14<sup>th</sup>, Feb., 2014.

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## 1   **Abstract**

2   We present here experimental results on aqueous-phase (A) photochemical oxidation (with UV and  
3   OH radicals generated from H<sub>2</sub>O<sub>2</sub> photolysis) and (B) direct photolysis (with only UV irradiation) of  
4   a methoxy-phenol, vanillin (VL), as a model compound from biomass burning. Both on-line aerosol  
5   mass spectrometric (AMS) characterization and off-line chemical analyses were performed. AMS  
6   analyses of dried atomized droplets of the bulk reacting mixtures showed that VL almost entirely  
7   evaporates during the drying process. Large amounts of organic mass remained in the particle phase  
8   after reactions under both conditions. Under condition (A), AMS measured organic mass first  
9   increased rapidly and then decreased, attributable to the formation of non-volatile products and  
10   subsequent formation of smaller and volatile products, respectively. The oxygen-to-carbon (O:C) ratio  
11   of the products reached 1.5 after about 80 minutes, but dropped substantially thereafter. In contrast,  
12   organic mass increased slowly under condition (B). The O:C ratio reached 1.0 after 180 minutes. In  
13   off-line analyses, small oxygenates were detected under condition (A), while hydroxylated products  
14   and dimers of VL were detected under condition (B). Particle hygroscopic growth factor (GF) and  
15   cloud condensation nuclei (CCN) activity of the reacting mixtures were found to depend on both  
16   organic volume fraction and the degree of oxygenation of organics. Results show that (1) aqueous-  
17   phase processes can lead to the retention of a large portion of the organic mass in the particle phase;  
18   (2) once retained, this portion of organic mass significantly changes the hygroscopicity and CCN  
19   activity of the aerosol particles; (3) intensive photochemical oxidation gave rise to an O:C ratio as high

1 as 1.5 but the ratio decreased as further oxidation led to smaller and more volatile products; and (4)  
2 polymerization occurred with direct photolysis, resulting in high-molecular-weight products of a  
3 yellowish color. This study demonstrates that aqueous-phase reactions of a methoxy-phenol can lead  
4 to substantial amount of secondary organic aerosol (SOA) formation. Given the vast amount of  
5 biomass burning input globally, model representation of either the SOA budget or their subsequent  
6 effects would not be adequate if the contribution of SOA formation from aqueous-phase reactions of  
7 methoxy-phenols is not considered.

## 1. Introduction

Reduction in saturation vapor pressure during chemical reactions, either by increasing the degree of oxygenation as in oxidation or increasing molecular weights as in polymerization (Kroll and Seinfeld, 2008), is the key to secondary organic aerosol (SOA) formation. Current models tend to underestimate the amount of SOA production (Heald et al., 2005). To close the gap, recent investigations have been searching for either missing precursors or mechanisms (Carlton et al., 2009) that would lead to products with substantially lower saturation vapor pressures than their precursors. Consideration of intermediately volatile organic compounds (IVOCs) as the missing SOA precursors (Donahue et al., 2006; Robinson et al., 2007) can possibly explain part of the underestimation. However, current studies have focused almost exclusively on the fossil-fuel-related long-chain alkanes in the IVOC range (Presto et al., 2010), which are precursors with no oxygen atoms in their structures. Some partially oxidized organic compounds whose saturation vapor pressures also fall in the IVOC range have not been adequately investigated. As for missing mechanisms, aqueous-phase oxidation has been advocated to be an important contributor to SOA in conditions where cloud, fog or wet aerosol particles are present (Blando and Turpin, 2000). There are numerous studies on aqueous-phase reactions but their focus has been limited to small aldehydes such as glyoxal and glycolaldehyde (Ervens et al., 2011), and recently pinene oxidation products (Lee et al., 2011).

Biomass burning is an important source of primary organic aerosol (POA) (de Gouw and Jimenez, 2009). It is also a source of SOA precursors (Cubison et al., 2011; Grieshop et al., 2009) because of the

1 numerous organic compounds with a wide range of volatilities emitted. Cellulose, hemi-cellulose, and  
2 lignin are major constituents of biomass, and biomass burning organic aerosol (BBOA) still carries  
3 some signatures linking to these phyto-macromolecules even after extensive thermal degradation.  
4 Levoglucosan from cellulose or hemi-cellulose and methoxy-phenolic compounds from lignin have  
5 been used as markers of BBOA for a long time (Simoneit, 2002). As an abundant BBOA component  
6 and a well-known BBOA marker, levoglucosan was believed to largely exist in the particle phase, but  
7 recent studies have suggested that it is semi-volatile, and its stability is being questioned because it  
8 can partition back to the gas phase for photochemical aging (Hennigan et al., 2010) and can also be  
9 oxidized in aqueous phase (Hoffmann et al., 2009). On the other hand, phenols and methoxy-phenols  
10 can contribute to 20-40% of particulate mass from burning hardwood and softwood (Hawthorne et al.,  
11 1989). Some methoxy-phenolic compounds will partition between the gas phase and the particle phase  
12 due to their relatively high saturation vapor pressures. For example, vanillin ( $C_8H_8O_3$ , 4-hydroxy-3-  
13 methoxybenzaldehyde, CAS#121-33-5) has a saturation vapor pressure of  $1.18 \times 10^{-4}$  mmHg at 298 K  
14 (Yaws, 1994), equivalent to a saturation concentration of approximately  $10^3 \mu\text{g}/\text{m}^3$ , and belongs to the  
15 IVOC range (Donahue et al., 2006). Its water solubility is 1.1 g/100 mL  $H_2O$  at 298 K (Budavari et al.,  
16 1996), which puts it in the slightly water soluble category. The physical Henry's law constant of  
17 vanillin is thus estimated to be  $4.65 \times 10^5$  M/atm at 298 K (Yaws, 1994), which is very close to the  
18 effective Henry's law constant of glyoxal (Ip et al., 2009). This means that vanillin, like glyoxal, will  
19 preferably be in the condensed phase if liquid water is available. In fact, it was observed that the

1 efficiency with which fog water scavenges methoxy-phenolic volatile compounds such as vanillin was  
2 higher than 90% (Collett Jr et al., 2008). Unlike glyoxal which self-polymerizes more favorably than  
3 evaporates, vanillin is driven to the gas phase upon water evaporation (see discussions later). Upon  
4 dilution in the atmosphere, the intermediately volatile or semi-volatile species such as vanillin in fine  
5 particles will at least partly evaporate, controlled presumably by absorptive partitioning. If coarse-  
6 mode droplets bearing a large amount of liquid water are available, the vapor-phase vanillin is then  
7 able to dissolve into the large droplets because of its slight solubility in water, resulting in a bimodal  
8 distribution. Indeed, this bimodal distribution of vanillin was clearly observed in a recent study (Wang  
9 et al., 2011).

10 We present experimental results of aqueous-phase reactions of a methoxy-phenolic model  
11 compound, vanillin, under two conditions: (A) UV (254 nm) + H<sub>2</sub>O<sub>2</sub> and (B) UV (254 nm) only. The  
12 aqueous bulk reacting mixture was partially atomized and dried for on-line analyses. The experimental  
13 setup, resembling a cloud-phase reaction/evaporation process, allowed us to characterize the evolution  
14 of organics in both the bulk aqueous solution and the particulate organics that remained after drying.  
15 The organics in the particles and in the solution during reactions were characterized by on-line aerosol  
16 mass spectrometry (AMS) and off-line analyses, respectively. Vanillin evaporated, as noted above, but  
17 large amounts of organic mass retained after reactions under both conditions (A) and (B), indicating  
18 the ability to retain slightly volatile organic species in the particle phase by aqueous-phase reactions.  
19 Substantial differences in product identities and reaction kinetics were observed between these two

conditions. Furthermore, the water uptake ability and cloud condensation activity were also monitored with a hygroscopic tandem differential mobility analyzer (HTDMA) and a cloud condensation nuclei counter (CCNc). The hygroscopicity of particles after reactions was significantly altered, and will be discussed in conjunction with the degree of oxygenation of particulate organics at different time of the reactions.

## 2. Experimental

### 2.1 Aqueous-phase reaction and control experiments

The reaction apparatus was modified from Lee et al. (2011) and is shown in Supporting Information Figure S1. A TSI 3076 constant output atomizer was modified to include: a) a low-pressure mercury UV (254 nm) lamp (Pen-Ray, UVP, Canada); b) a circulating sample collection system; c) a magnetic stirrer; and d) a water bath with an aluminum foil cover to keep out natural light during the experiments. A volume of 300 mL aqueous solution of 0.1 mM ammonium sulfate (AS, Sigma,  $\geq 99.0\%$ ) and 0.1 mM vanillin (VL, Sigma-Aldrich, 99%), with or without 11.8 mM  $\text{H}_2\text{O}_2$  (Sigma-Aldrich, 34.5-36.5 wt%), was continuously atomized with compressed air to generate particles, which passed through a diffusion drier (BMI, Haywood, CA, USA) before going into the AMS, HTDMA and CCNc. The conditions and measurements in control (C\_1 to C\_4) and actual experiments (A\_1 to A\_9 and B\_1 to B\_7) are shown in Table 1. In some experiments (A\_9 and B\_7), aqueous solution samples were collected at different time intervals for off-line analyses. In the control

1 experiments with AS only and the UV lamp on (C\_1), AS and H<sub>2</sub>O<sub>2</sub> with the UV lamp on (C\_2), AS  
2 and VL without UV (C\_3), and AS, VL and H<sub>2</sub>O<sub>2</sub> without UV (C\_4), the mass fractions of organics  
3 were normally less than 5% after drying, with the rest of particle mass being AS. The concentrations  
4 of sulfate, ammonium and organics, as well as the ratios of organics to sulfate in those control  
5 experiments without any one of the critical elements (UV, H<sub>2</sub>O<sub>2</sub>, or VL) of the reactions are showed in  
6 Figure S2. Results from experiments C\_1 and C\_2 showed that the mass contribution from the  
7 background (water) was negligible. Results from experiment C\_3 indicated that VL evaporated almost  
8 completely during drying if no reaction occurred and those from experiment C\_4 indicated that there  
9 was little retainable product formed between VL and H<sub>2</sub>O<sub>2</sub>, or the reaction between them is too slow  
10 for the time scale of the experiments (3 hours).

## 11 2.2 AMS analysis

12 The high-resolution time-of-flight aerosol mass spectrometer (HR-ToF-AMS) has been described  
13 in detail elsewhere (DeCarlo et al., 2006). Briefly, particles leaving the diffusion dryer enter a set of  
14 aerodynamic lens for particle focusing and a particle time-of-flight (pToF) chamber for sizing. They  
15 then arrive at a heater with a temperature of 600 °C for vaporization. Vapors of the non-refractory (NR)  
16 components are then ionized by electrons emitted by a tungsten filament. Ions are then transmitted to  
17 an ion time-of-flight (iToF) chamber for mass analysis and finally detected by a multi-channel plate  
18 (MCP) detector. In all the experiments, the HR-ToF-AMS was operated under two modes, each for 2  
19 minutes. The first mode was operated under a V-shaped iToF (for lower mass resolving power but with



1 higher sensitivity) plus pToF (for sizing). The second mode was operated under a W-shaped iToF (for  
2 higher mass resolving power but with lower sensitivity). The AMS data were analyzed using two  
3 toolkits, SQUIRREL and PIKA (Sueper, 2012), which are based on Igor Pro (Wavemetrics, Lake  
4 Oswego, OR). Based on previous studies (Li et al., 2011b; Zhang et al., 2005), the fragmentation table  
5 was modified for  $m/z$  18 (or  $\text{H}_2\text{O}^+$  ion in W mode) and  $m/z$  28 (or  $\text{CO}^+$  ion in the W mode) for organics.  
6 Details of the modification can be found in the Supporting Information (Section 3.1, Figure S4 and  
7 Table S1). Background mass spectra from control experiment C\_1 (Section 3.2 in Supporting  
8 Information) were also subtracted from mass spectra of experiments to obtain representative AMS  
9 mass spectra of the particle-phase organics formed during reactions (see Section 3.1 below). The AMS  
10 measures organics, sulfate, ammonium, nitrate and chloride. Nitrate and chloride in our experiments  
11 were negligible as expected. We do not exclude the possibility of organic sulfate formation during our  
12 experiments. However, due to the high temperature (600 °C) and high energy ionization, even if they  
13 were present, organic sulfates were likely fragmented to give sulfur-containing ions that are similar to  
14 inorganic sulfate, and was mostly counted as inorganic sulfate contribution. Therefore, the sulfate  
15 content is considered to be constant in the bulk solution during the experiments and will be used to  
16 normalize the measured organic content (Lee et al., 2012). Any fluctuations in the measured sulfate  
17 content were assumed to be due to the fluctuations in particle generation and the difference in  
18 collection efficiency (CE) as a result of the different organic contents of particles (Middlebrook et al.,  
19 2012).

### 2.3 OH concentration estimation

Under condition (A), H<sub>2</sub>O<sub>2</sub> photolysis was used to generate OH radicals. In the H<sub>2</sub>O<sub>2</sub> decay experiments, H<sub>2</sub>O<sub>2</sub> concentrations were measured at different time intervals using a colorimetric method (Allen et al., 1952). A typical calibration curve is shown in Figure S5. The decay rate ( $k_{\text{obs}}$ ) of H<sub>2</sub>O<sub>2</sub> was determined to be  $1.9 \times 10^{-4} \text{ s}^{-1}$  (Figure S6), similar to those reported in a previous study (An et al., 2001) that used the same brand and model of UV lamp. A photolysis rate constant of  $1.0 \times 10^{-4} \text{ s}^{-1}$  was determined by taking into account reactions with other reactive oxygen species such as OH and HO<sub>2</sub> radicals (see Supporting Information Section 4, Table S2 and Figure S6 for details). The OH concentration in the aqueous solution was estimated by two approaches (see Supporting Information Section 4 and Figures S7 and S8 for details): (1) assuming a pseudo-steady-state of OH radicals, and (2) solving the stiff ordinary differential equations with no constraint on the OH radical concentration. The first approach resulted in an OH concentration of about  $7.2 \times 10^{-12} \text{ M}$ , while the second one gave a concentration of about  $7.0 \times 10^{-12} \text{ M}$ . The OH concentration in the aqueous solution is thus believed to be one order of magnitude higher than that in cloud water (Figure S7), which is generally in the order of  $10^{-13} \text{ M}$  (Ervens et al., 2013; Warneck, 2003), but is close to that in fog or wet aerosol particles (Ervens et al., 2011).

### 2.4 Off-line analyses

Aqueous-phase samples collected at different time intervals were analyzed using three off-line

1 techniques. An ultra-performance liquid chromatograph (UPLC) with a diode array detector (DAD)  
2 was used to quantify the reactant VL and an expected product, vanillic acid (VA), in experiments A\_1  
3 to A\_3 and B\_1 to B\_3. An UPLC with electrospray ionization (ESI) coupled with a time-of-flight  
4 mass spectrometric detector (ToF-MS) was used to identify reaction products that have relatively large  
5 molecular weights, i.e., mainly those from condition (B). Product identification with derivatizations  
6 prior to gas chromatography and mass spectrometry (GC-MS) was also conducted for small  
7 oxygenated products, which were mainly found under condition (A). Two derivatization methods were  
8 employed. The first one uses the  $\text{BF}_3$ /butanol to convert carboxylic acids to butyl esters (Li and Yu,  
9 2005). The second one uses pentafluorophenylhydrazine (PFPH) to convert carbonyl compounds to  
10 their PFPH derivatives (Dron et al., 2008). Details of off-line analyses can be found in the Supporting  
11 Information Section 5.

## 12 *2.5 HTDMA and CCNc analyses*

13 During experiments A\_4 to A\_8 and B\_1 to B\_6, particles leaving the diffusion dryer were  
14 characterized by an HTDMA system (BMI, Haywood, CA, USA). Dried particles entered a differential  
15 mobility analyzer (DMA1) operating under dry conditions (relative humidity,  $\text{RH} < 5\%$ ) and left  
16 DMA1 with a mono-disperse particle size with mobility diameter,  $D_{\text{dry}} = 100 \text{ nm}$ . They then entered a  
17 humidifier with a controlled RH of 90% to effect hygroscopic growth. After growth, the size  
18 distribution over the mobility diameter  $D_{90}$  was measured with a second DMA (DMA2) operating in  
19 scanning mode and coupled with a mixing condensation particle counter (MCPC). The hygroscopic

1 growth factor,  $GF_{90}$ , defined as the ratio of the humidified (at 90% RH) particle diameter to the dry  
2 particle diameter, was then obtained. Note that under condition (A) when substantial amounts of small  
3 oxygenates were formed, the dried particles may not be completely water free. Although there is  
4 uncertainty in the estimation without the exact knowledge of the amounts of all the small oxygenates  
5 and high electrolyte concentrations, EIM-II (Clegg et al., 1998) predicted a <2% difference in the  $D_{dry}$ .

$$GF_{90} = \frac{D_{90}}{D_{dry}}$$

7 In the CCN measurements (A\_7 to A\_8 and B\_4 to B\_6) for the cloud condensation activity, dry  
8 particles leaving the diffusion dryer passed through a TSI (Shoreview, MN) DMA and were split into  
9 two streams. One stream was directed to a TSI water-based CPC and the other to the CCN counter  
10 (CCNc, model CCN-200, DMT, Boulder, CO). The particle number concentration ( $N_{CN}$ ) was obtained  
11 by the CPC. In the CCNc, an effective water vapor supersaturation (ss) of 0.1% was set. Particles that  
12 had a critical supersaturation below the set value were then activated and grew into droplets. The  
13 concentration ( $N_{CCN}$ ) and size distribution of activated droplets were detected by an optical particle  
14 counter at the bottom exit of the CCNc column.

15 The CCN activation of the particles was examined by plotting the CCN efficiency against the  
16 diameter ( $D_p$ ) of dry particles selected by a DMA and counted by a CPC. The CCN efficiency is defined  
17 as the ratio of the concentration of the activated droplets to that of the dry particles ( $N_{CCN}/N_{CN}$ ). After  
18 making a correction for doubly charged particles, each CCN efficiency spectrum was fitted with a

cumulative Gaussian distribution function using a non-linear least-square fitting routine. The  $D_{50}$ , indicating 50% of the particles at this diameter are activated, was then obtained. Since the supersaturation was fixed during the experiment, changes in  $D_{50}$  would reflect changes in the CCN activity of the particles, which is related to the particle chemical composition.

### 3. Results and discussion

#### 3.1 AMS Organic fractions and mass spectra

Aqueous-phase reactions under both conditions (A) and (B) were triggered by UV light. Under condition (A) when  $H_2O_2$  was present, OH radicals were generated and oxidation of VL proceeded mainly via reactions with OH. Under condition (B) when  $H_2O_2$  was absent, direct photolysis of VL was the main reaction route (Benitez et al., 1997), although a small amount of reactive oxygen species including  $H_2O_2$  was very likely generated (Anastasio et al., 1997). The differences between these two conditions can be clearly seen from the increases in AMS-measured particulate organic mass concentration (green line) and the ratio of organic to sulfate mass ( $Org/SO_4^{2-}$ ) (black circles) in Figure S2 (for experiments A\_1 to A\_3 in the upper panel and B\_1 to B\_3 in the lower panel). Shown in Figure 1 are the 10-minute running average  $Org/SO_4^{2-}$  ratios for all the experiments (A\_1 to A\_8 in panel A, and B\_1 to B\_6 in panel B), together with the average mass fractions of organics (green) and ammonium sulfate (red). Under condition (A), the  $Org/SO_4^{2-}$  ratio increased rapidly and reached 1.58 (the original  $VL/SO_4^{2-}$  mass ratio of the reacting solution) in around 30 minutes. The overshooting of

1 the Org/SO<sub>4</sub><sup>2-</sup> ratio up to 1.80 was probably because addition of oxygen mass due to functionalization  
2 out-weighed the loss of VL and/or volatile products during the photochemical oxidation. Note that the  
3 usage of the default relative ionization efficiency of organics (1.4) may introduce an uncertainty of  
4 ~20% and ~15% in the estimation of the organic and the sulfate concentrations (Bahreini et al., 2009),  
5 respectively, which would result by error propagation in an uncertainty of ~25% for the highest  
6 Org/SO<sub>4</sub><sup>2-</sup> ratio. After around 40 minutes, the Org/SO<sub>4</sub><sup>2-</sup> ratio started to decrease slowly. After around  
7 3 hours, the Org/SO<sub>4</sub><sup>2-</sup> ratio fell to around 0.1, indicating that the organics were mostly converted to  
8 volatile products by extensive fragmentation.

9 Under condition (B), the Org/SO<sub>4</sub><sup>2-</sup> ratio increased steadily throughout the 3-hour reaction period.  
10 It reached around 1.0 and the organic mass fraction was around 40% by the end of the experiments.  
11 Hence, without a high OH radical concentration, direct photolysis of VL by UV light with a wavelength  
12 of 254 nm for 3 hours resulted in the retention of only approximately 60% (the highest Org/ SO<sub>4</sub><sup>2-</sup> ratio  
13 of around 1.0 divided by original VL/ SO<sub>4</sub><sup>2-</sup> ratio of 1.58) of the organic mass in the particle phase  
14 after reactions and drying. The rest is likely unreacted VL or its volatile products, both of which would  
15 evaporate during the drying process.

16 In the high-resolution mass spectra obtained under the W mode (Figure 2A), ions with high *m/z*  
17 values were present in the first 10 minutes of reactions under condition (A). For example, the peak of  
18 *m/z* 137 (see grey arrow in panel A) was dominated by an ion with a formula of C<sub>7</sub>H<sub>5</sub>O<sub>3</sub><sup>+</sup>. This *m/z*  
19 value has been observed in biomass burning organic aerosols and was assigned to lignin-related ions

1 with methoxy-phenolic structures (Li et al., 2011b). Therefore, this ion is believed to be a fragment of  
2 the ring-retaining products that still possess the methoxy-phenolic structure. After 40 minutes, the ions  
3 with high  $m/z$  values became less abundant. Instead, the ions of  $\text{CO}^+$  ( $m/z$  28) and  $\text{CO}_2^+$  ( $m/z$  44)  
4 dominated. At 120 minutes, these two ions became overwhelmingly dominant in the mass spectrum.  
5 In contrast, the ions with high  $m/z$  values were still of relatively high fractions under condition (B)  
6 throughout the 3-hour experiment (Figure 2B). The peak at  $m/z$  137 remained observable even at 120  
7 minutes.

8       The differences in the evolution of mass spectral features under these two conditions indicate that  
9 the retainable products observed were substantially different within the experimental time scale of 3  
10 hours. The disappearance of high-molecular-weight ions and emergence of ions representing small  
11 oxygenates (e.g.,  $\text{CO}^+$  and  $\text{CO}_2^+$ ) under condition (A) suggested that fragmentation might dominate  
12 over functionalization. Fragmentation was very likely via ring-opening pathways, leading to the  
13 formation of carboxylic acids which generated high intensity of  $m/z$  44 ( $\text{CO}_2^+$ ) in AMS mass spectra.  
14 Under condition (B), on the other hand, the preservation of the methoxy-phenolic structure was  
15 evidenced by the presence of  $m/z$  137 ( $\text{C}_7\text{H}_5\text{O}_3^+$ ). Reactions might have proceeded mainly via  
16 functionalization with the ring retained. Note that even under condition (B), the intensities of high-  
17 molecular-weight ions decreased but those of  $m/z$  18 and  $m/z$  44 increased too, which suggested that  
18 fragmentation has occurred, albeit at a less extent than that under condition (A).

### 19 3.2 Degree of oxygenation

1       The oxygen-to-carbon (O:C) and hydrogen-to-carbon (H:C) ratios are shown in Figure 3 for both  
2 conditions. The O:C ratio under condition (A) increased for the first 80 minutes before it decreased.  
3 This appears to contradict the extensively oxidative environment during the experiments. However,  
4 since the AMS-measured organic products were only those that remained in the particle phase, some  
5 of the more oxygenated products, for example, formic acid ( $\text{O:C} = 2$ ) from fragmentation reactions, as  
6 a very oxygenated and volatile product of the oxidation, may have escaped from the particle phase and  
7 evaded detection by the AMS. A similar fragmentation-induced formation of “less oxygenated” SOA  
8 was also observed with a decrease in O:C during the gas-phase photochemical oxidation of biomass  
9 burning organic aerosols (Heringa et al., 2011). On the other hand, mineralization to  $\text{CO}_2$  as the end  
10 product may also have caused a substantial decrease in the O:C of the particulate organics measured.  
11 Although not measured,  $\text{CO}_2$  very likely formed during the experiment because the O:C at 80 minutes  
12 had already increased to 1.5, approaching the upper limit for the commonly observed particulate  
13 organic species (e.g., 2.0 for oxalic acid). The O:C ratio under condition (B) increased, first rapidly  
14 and then slowly during the 3 hours of reactions. This indicates that oxidation under condition (B) was  
15 not as complete as that under condition (A), because O:C under condition (B) was only  $\sim 1.0$  at the end  
16 of the experiments. Therefore, the organic products that retained in the particle phase at the time of  
17 highest O:C ratio under condition (B) ( $\sim 3$  hours) are believed to be less oxygenated than those under  
18 condition (A) ( $\sim 80$  min). The H:C ratio decreased slightly under both conditions at the beginning of  
19 the experiments. While the H:C ratio was relatively stable throughout the rest of the experiment under



condition (B), it increased slightly under condition (A) near the end of the experiment.

### 3.3 Kinetics of vanillin decay

The reactant VL and an expected product vanillic acid (VA) were quantified using standard compounds and offline UPLC-DAD (see Supporting Information Section 5.1, Figure S9 and Figure S10). Figures 4A and 4B show the ratios of these two compounds (green solid circles and red solid triangles) to sulfate (assumed constant at 0.1 mM) during experiments under conditions (A) and (B), respectively. Under condition (A), VL/SO<sub>4</sub><sup>2-</sup> decreased rapidly during the first 30 minutes after the UV lamp was switched on, coinciding with the rapid increase in Org/SO<sub>4</sub><sup>2-</sup> revealed by the AMS measurements. VA/SO<sub>4</sub><sup>2-</sup> increased slightly in the first 10 minutes but decreased to almost zero after 30 minutes, suggesting that VA was only an important product at the beginning of the experiment under condition (A) but was almost completely converted to other products after 30 minutes of photochemical oxidation. Under condition (B), VL/SO<sub>4</sub><sup>2-</sup> decreased slowly, corresponding to a similarly slow increase in Org/SO<sub>4</sub><sup>2-</sup> observed in the AMS measurements. VA remained as a minor product during the reactions under condition (B), suggesting some other products (see below) were formed during the consumption of VL.

Figures 4C and 4D show the decay rate ( $k_{\text{obs}}$ , s<sup>-1</sup>) of VL under conditions (A) and (B), respectively. The decay rate under condition (A) was one order of magnitude larger than that under condition (B). Using the estimated OH steady-state concentration ( $7.0 \times 10^{-12}$  M), a bimolecular reaction rate constant

of VL and OH under condition (A) was estimated to be  $4 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$  (see Supporting Information Section 4 and Figure S8 for details). For condition (B), reactions were likely triggered by activation of the aromatic VL by UV light, followed by production of a smaller amount of  $\text{H}_2\text{O}_2$  (Anastasio et al., 1997) than under condition (A), which in turn generated a much lower concentration of OH radicals than under condition (A). In the initial stage of the experiment (i.e. the first 60 minutes), the decay rate was determined to be  $2.3 \times 10^{-4} \text{ s}^{-1}$ . This decay rate is close to that ( $3.6 \times 10^{-4} \text{ s}^{-1}$ ) in Benitez et al. (1997), although polychromatic light (strongest line at 366 nm) was used in that study.

### *3.4 Other reaction products*

Mass spectra of some of the products identified by UPLC-ESI-ToF-MS and GC-MS are shown in Figure 5 and Figure 6, respectively. Extractive ion chromatograms from UPLC-ESI-ToF-MS are shown in Figure S11. All mass spectra are shown in the Supporting Information in Figure S12 (for UPLC-ESI-ToF-MS) and Figure S13 (for GC-MS).

Using the UPLC-ESI-ToF-MS, no major products were detected in samples under condition (A), probably due to poor separation in LC and inefficient ionization by ESI for those products formed under condition (A). As discussed earlier in Section 3.1, ring-opening and highly oxygenated products with low molecular weights (e.g., small carboxylic acids or aldehydes) are expected to form under extensive photochemical oxidation under condition (A). Indeed, it is not efficient to separate small oxygenates in a reverse-phase LC column (Buszewski and Noga, 2012) and to ionize small oxygenates

1 that are too hydrophilic by ESI (Cech et al., 2001). Multiple products were observed by the UPLC-  
2 ToF-MS in samples under condition (B), and they are listed in Table S3. A number of monomeric  
3 products with retention times of around 10 minutes (Figure S11), including B168\_a (Figure 5a),  
4 B168\_b, and B168\_c with a molecular formula of  $C_8H_8O_4$ , as well as B184\_a and B184\_b with a  
5 molecular formula of  $C_8H_8O_5$ , were observed after 40 minutes of reaction under condition (B). Among  
6 them, B168\_b is believed to be VA as detected by the UPLC-DAD method, based on its fragmentation  
7 pattern which showed a  $CO_2$  loss (Li et al., 2011a). Similarly, the observed product B184\_a (Figure  
8 5b) is also believed to bear a carboxylic group based on its fragmentation pattern which showed a  $CO_2$   
9 loss (Table S3). Other isomers might be di- or tri-hydroxylated products with one or two more hydroxyl  
10 groups (in addition to the original one) at the three available positions on the aromatic ring.

11 Seven products with longer retention times (32-35 minutes) were observed after 120 minutes of  
12 reaction under condition (B), as shown in Table S3 and Figure S11. The mass spectra of these products  
13 are shown in Figure S12. The oligomeric product B302\_a (Figure 5c) is believed to be a dimer formed  
14 via radical-radical polymerization, similar to reactions of other phenols (Sun et al., 2010). Mass spectra  
15 of other observed oligomers contain less information on fragmentation for further structural  
16 elucidation thus the structures of those products are not proposed and their formation mechanisms  
17 remain unknown. Nevertheless, those products have similarly high conjugation as B302\_a because  
18 their double bond equivalency is higher than (or at least close to) that of B302\_a (Table S3).  
19 Additionally, a yellowish color was observed for solutions after experiments under condition (B) while

1 it was not observed under condition (A). This observation suggests that the highly conjugated products  
2 as detected by the UPLC-ToF-MS might be responsible for the light absorption of the solution.

3 From samples under condition (A), four small carboxylic acids including glyoxylic acid, oxalic  
4 acid, malonic acid, and pyruvic acid, as well as two aldehydes including formaldehyde and glyoxal,  
5 were identified using the two aforementioned derivatization methods prior to GC-MS analyses (Figure  
6 S13). Figure 6 shows the standard and sample mass spectra of derivatized oxalic acid (Figure 6a),  
7 glyoxylic acid (Figure 6b), and glyoxal (Figure 6c). These small oxygenates were detected in samples  
8 at 80-120 minutes after the start of the reactions. The formation of these small oxygenates confirms  
9 that with intensive photochemical oxidation under condition (A), vanillin molecules were extensively  
10 fragmented (ring-opening) to yield smaller and more functionalized products.

11 From the above product information, it can be concluded that the reaction pathways under  
12 conditions (A) and (B) are very different. Under condition (A), photochemical oxidation may proceed  
13 via ring-opening pathways that lead mainly to small oxygenates, resulting in a very high degree of  
14 oxygenation as indicated by the O:C ratios seen in the AMS results (Figure 3A). Under condition (B),  
15 reactions seemed to occur mostly via ring-retaining pathways, resulting in the addition of one or more  
16 hydroxyl groups. Dimerization of vanillin also occurred under this condition. These two reaction  
17 pathways under condition (B) resulted in a mild increase in the O:C ratios, consistent with the AMS  
18 measurements (Figure 3B). However, all the products tentatively identified by the UPLC-ToF-MS  
19 have O:C ratios far below 0.9 at 120 minutes, the O:C ratio obtained from AMS measurements. The

products observed by the UPLC-ToF-MS represented those in the solution, and some of them (with low O:C ratios) might not retain in the particles. Those that retained (but not identified using the current techniques) might have higher O:C ratios than those observed by the UPLC-ToF-MS method. It is possible that some small oxygenates with high O:C ratios were also produced under condition (B), albeit to a lesser extent, as the UPLC-DAD measurements indicated (see Supporting Information Section 5.1).

### *3.5 Hygroscopicity and CCN activity*

The mixed inorganic (AS) and organic particles were characterized by HTDMA and CCNc analyses during the reactions under both conditions. AS had a measured (Figures 7A and 7C, before reactions)  $GF_{90}$  of 1.68 for a dry size of 100 nm, close to that in the literature (Duplissy et al., 2009). A critical activation diameter of ~115 nm for AS at  $ss = 0.1\%$  was measured (Figures 7B and 7D, before reactions), slightly smaller than the 129 nm measured at 0.09 %  $ss$  (Roberts and Nenes, 2005). Under condition (A), the  $GF_{90}$  decreased rapidly as the organic volume fraction increased. Interestingly, the lowest GF did not correspond to the highest organic volume fraction (Figure 7A). Instead, the GF started to increase slightly as the organic volume fraction increased from around 0.55 to 0.60 (AS volume fraction from 0.45 to 0.40), corresponding to the reaction times of around 30 minutes and 40 minutes, respectively. This indicates that at 40 minutes, the organic products were more hygroscopic than those at 30 minutes, assuming the contribution from the inorganic component (AS) was constant. This is in agreement with the observation that the degree of oxygenation, as indicated by the O:C ratio,

1 increased rapidly in the first 40 minutes of the experiments under condition (A). After 40 minutes, the  
2 organic volume fraction started to decrease and the  $GF_{90}$  further increased. After 80 minutes, the GF  
3 continued to increase although the O:C ratio of organics began to decrease. This is because the organic  
4 volume fraction had decreased substantially and the  $GF_{90}$  was overwhelmingly dominated by the  
5 inorganic AS again. The  $D_{50}$  showed a reverse trend in Figure 7B. Under condition (B), however, the  
6 GF decreased continuously, with the rate of decrease higher at the beginning than at the end of the  
7 experiments (Figure 7C). Correspondingly, the  $D_{50}$  increased throughout the experiment under  
8 condition (B) (Figure 7D).

9       To quantitatively understand the effects of organic products from these two conditions on particle  
10 hygroscopicity, the hygroscopic growth factors of the organic species at different reaction time  
11 intervals were estimated using the Zdanovskii-Stokes-Robinson mixing rule. The details of  
12 hygroscopic growth factor estimation are shown in Section 6 of the Supporting Information. The  
13 hygroscopic growth factors of organic products under the two conditions as a function of the O:C ratio  
14 are shown in Figure 8. Only data points with organic mass fraction  $> 30\%$  (volume fraction  $> \sim 35\%$ )  
15 were included in this plot to avoid overwhelmingly dominant contribution by AS. Also shown in  
16 Figure 8 is the fitted line from the literature (Massoli et al., 2010). A linear dependence of  $GF_{org}$  on the  
17 O:C ratio was observed in our experiments under both conditions. But the slopes and the intercepts are  
18 quite different from those in Massoli et al. (2010). Firstly, for the OH oxidation experiments under  
19 condition (A), the slope is much smaller than that in Massoli et al. (2010). This may be because the

1 range of O:C values from our experiments in condition (A) is much higher (generally > 0.5 for data  
2 points with organic mass fraction > 30%) than those from chamber or field studies as in Massoli et al.  
3 (2010). The data points in Massoli et al. (2010) also started to level off with O:C > 0.8. Indeed, if the  
4 O:C ratio keeps increasing up to 2.0 (e.g., for oxalic acid), neither of these two linear relationships  
5 would hold because the GF<sub>90</sub> values of smaller carboxylic acids such as oxalic acid are generally less  
6 than 1.5 (Peng et al., 2001). For experiments conducted under condition (B), the GF<sub>org</sub> was  
7 substantially lower than in the Massoli et al. (2010) and under condition (A) in our study. This indicates  
8 that the relationship between GF and O:C ratio of those highly conjugated organic compounds such as  
9 those formed from condition (B) and generally present in BBOA may be substantially different from  
10 that for chamber SOA.

### 11 *3.6 Aqueous-phase processing*

12 To relate the current experiments to ambient aqueous droplets, we now compare the time scales  
13 of these experiments to the atmospheric life times of clouds and VL. The OH concentration in clouds  
14 is usually in the order of 10<sup>-13</sup> M (Ervens et al., 2011;Warneck, 2003), while the OH concentration in  
15 our experiments was one order of magnitude higher. With a concentration of 1.84 × 10<sup>-13</sup> M (Warneck,  
16 2003), the time scale equivalent to our experimental time (3 hours) would be ~110 hours for a similar  
17 OH exposure under ambient conditions. Given a cloud condensation-evaporation cycle of 0.5 – 3 hours  
18 and ~10 cycles for most cloud water nuclei (Warneck, 2000), our experimental time scale for condition  
19 (A) was a few times that of the typical cloud life time. Therefore, the first 40-60 minutes of the

1 experiments under condition (A) would be more relevant to ambient clouds. Oxidation in cloud is  
2 effective when biomass burning aerosols are activated to form cloud droplets, where semi-volatile  
3 methoxy-phenolic compounds can be dissolved into the aqueous phase and a reasonably high  
4 concentration of OH can be accommodated by the droplets. Between the cloud formation cycles, water  
5 evaporates and only aerosol particles remain (together with oxidation products from the previous cloud  
6 cycle). In that case, oxidation reactions in the gas phase may be more important than in the particle  
7 phase since vanillin exists preferably in the gas phase.

8 For condition (B), on the other hand, oxidation reaction is not affected by the availability of OH  
9 from direct photolysis of H<sub>2</sub>O<sub>2</sub>. However, it is very likely that trace amounts of H<sub>2</sub>O<sub>2</sub> would be  
10 generated during the direct photolysis of VL, as is the case during the direct photolysis of other  
11 phenolic compounds (Anastasio et al., 1997). This small amount of H<sub>2</sub>O<sub>2</sub> can also form OH radicals,  
12 albeit to a much lower concentration than under condition (A). The measured decay rate of  $2.3 \times 10^{-4}$   
13 s<sup>-1</sup> is close to the loss rates of common water-soluble organic compounds due to aqueous-phase  
14 processes under ambient conditions (Ervens et al., 2013). This decay rate is at least comparable to the  
15 gas-phase loss rate ( $1 \times 10^{-4}$  s<sup>-1</sup>) for VL, assuming a gas-phase oxidation rate constant of  $10 \times 10^{-11}$   
16 cm<sup>3</sup> molecules<sup>-1</sup> s<sup>-1</sup> (Coeur-Tourneur et al., 2010) and a gas-phase OH concentration of  $1 \times 10^6$   
17 molecules cm<sup>-3</sup> (Ervens et al., 2013). Note that this decay rate was measured with a UV lamp of 254  
18 nm, lower than wavelength with significant actinic flux in the troposphere (>290 nm). A simple  
19 calculation shows that with the absorption cross section of VL estimated from its molar absorption



coefficient and an assumed quantum yield of 0.28 (resulting in the measured decay rate at 254 nm), the photolysis rate in the wavelength range of tropospheric importance (295-425 nm) is  $7.06 \times 10^{-4} \text{ s}^{-1}$ . The details of the calculation are provided in Section 7 in the Supporting Information. Therefore, even without H<sub>2</sub>O<sub>2</sub> in the aqueous phase, the loss of methoxy-phenolic compounds like VL through the aqueous-phase process (with UV light) can be as important as that through gas-phase oxidation.

#### 4. Atmospheric implications

The current work demonstrates that both (A) photochemical oxidation and (B) direct photolysis in aqueous phase can convert the intermediately volatile and slightly water-soluble, vanillin into less volatile products. Unlike the precursor vanillin, some of these products can remain after the drying process and be detected as particulate organics. The products from photochemical oxidation and direct photolysis are different. Significant changes in the hygroscopicity and CCN activity of the particles were observed in our experiments using AS and VL as initial components. The effects of organic products on aerosol hygroscopicity under the two conditions investigated are different, in line with their different structural features. Two interesting phenomena were observed. First, under condition (A), the degree of oxygenation first increased and then decreased. Second, a yellowish coloration of the reacting mixtures and high-molecular-weight products was observed mainly under condition (B). The first phenomenon suggests that aqueous-phase oxidation as under condition (A) may also have a competition between functionalization and fragmentation (Lee et al., 2012). The second phenomenon suggests that OH oxidation under condition (A) may lead to extensive carbon backbone breakage,

1 while direct photolysis under condition (B) may induce radical-radical polymerization (Sun et al.,  
2 2010). Hydroxylated products and dimers (e.g., B302\_a) are likely responsible for the yellowish color  
3 observed under condition (B). This observation suggests that direct photolysis of methoxy-phenolic  
4 compounds could contribute to the light-absorbing “brown carbon” observed in ambient aerosol  
5 samples (Andreae and Gelencser, 2006), although differences between actinic radiation in typical  
6 tropospheric environments ( $> 290$  nm) and the short UV wavelength (254 nm) used in the experiments  
7 need to be considered. During one of the trial experiments, the same yellowish color was also observed  
8 under condition (A) at the same  $\text{H}_2\text{O}_2$  concentration (11.8 mM) but a 10 times higher VL concentration  
9 (1 mM) than in all other experiments. The formation of the yellowish color may thus be dependent on  
10 the concentration ratio between VL and OH radicals. If an excess amount of VL is available,  
11 hydroxylated products and dimers with a yellowish color can still be formed under condition (A).

12 Well known as a source of POA, biomass burning also contributes to SOA formation by producing  
13 a large amount of slightly volatile and partially oxygenated species. Among many types of organics  
14 that are emitted from biomass burning, the methoxy-phenolic compounds of various levels of volatility  
15 deserve more attention regarding their contribution to SOA formation. It has been shown that they can  
16 be oxidized in gas phase to produce less volatile products, forming SOA (Cubison et al., 2011; Grieshop  
17 et al., 2009; Yee et al., 2013; Liu et al., 2012). Early studies using phenols as model compounds have  
18 also shown the potential role of this broad class of partially oxygenated compounds in SOA formation  
19 through aqueous-phase oxidation (Sun et al., 2010; Chang and Thompson, 2010). This study further

demonstrates that the potential contribution of aqueous-phase oxidation, either photochemical oxidation or direct photolysis, to SOA formation is quite high. Once formed and mixed with other species such as inorganic salts in the particle phase, the SOA from biomass burning can strongly affect the physical properties, such as hygroscopicity and CCN activity, of atmospheric aerosol particles.

It should be noted that the current study only used a single model compound of methoxy-phenols with a relatively high concentration (0.1 mM) in aqueous phase to demonstrate the reactivity. Other methoxy-phenols may vary in (i) the number of methoxy groups (e.g., p-hydroxyl, guaiacyl, and syringyl have 0, 1, and 2 methoxy groups, respectively), (ii) the number of carbon left in the propyl chain (for instance, VL only has one carbon left), and (iii) the type and position of functional groups in the propyl chain (OH, C=O, or COOH in  $\alpha$ ,  $\beta$ , or  $\gamma$  position). These methoxy-phenols may vary significantly in both saturation vapor pressure and water solubility. Although their reactivity is expected to be similar to that of VL because they share similar aromatic and partially oxygenated characteristics, their actual contribution to SOA during aqueous-phase processing warrants further investigation. On the other hand, pH in the aqueous droplets will have an effect on the aqueous-phase reactions (Ervens et al., 2011). Commonly found acidic aqueous phase in the atmosphere may not affect the production of OH radical from H<sub>2</sub>O<sub>2</sub> photolysis as shown in Figure S6, where ammonium bisulfate was used as the medium for H<sub>2</sub>O<sub>2</sub> UV photolysis at 254 nm, with negligible difference as compared to when AS solution was used as the medium. But the acidic environment may promote hydration of some carbonyl compounds, thus alter the kinetics and even branching ratios of aqueous-

phase reactions. In the current study, the pH of the aqueous solution was not controlled, and a pH value was approximately 6 with 0.1 mM of AS. Further investigation is needed to ascertain the role of acidity in the aqueous-phase reactions of methoxy-phenolic compounds. Moreover, the OH radical concentration used was also one order of magnitude higher than typical ambient concentrations. Nevertheless, the current work highlights the importance of SOA formation and its subsequent effects on aerosol particles from the aqueous-phase processing of methoxy-phenolic compounds generated by biomass burning.

## **5. Acknowledgements**

This work was supported by the Research Grants Council of the Hong Kong Special Administrative Region, China (project no. 600909 and 600413). The authors are very grateful to Mr. W.H. Fan, Dr. L.M. Yang, and Prof. L.E. Yu of the National University of Singapore for performing the UPLC-ToF-MS analyses.

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## 7. Table

Table 1. Summary of experimental conditions and measurements.

Expt #	Condition				Off-line <sup>e</sup>	AMS <sup>f</sup>	HTDMA <sup>g</sup>	CCN <sup>c</sup> <sup>h</sup>
A_1	AS <sup>a</sup>	VL <sup>b</sup>	H <sub>2</sub> O <sub>2</sub> <sup>c</sup>	UV on <sup>d</sup>	Y	Y		
A_2	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on	Y	Y		
A_3	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on	Y	Y		
B_1	AS	VL	–	UV on	Y	Y	Y	
B_2	AS	VL	–	UV on	Y	Y	Y	
B_3	AS	VL	–	UV on	Y	Y	Y	
A_4	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	
A_5	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	
A_6	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	
A_7	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	Y
A_8	AS	VL	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	Y
B_4	AS	VL	–	UV on		Y	Y	Y
B_5	AS	VL	–	UV on		Y	Y	Y
B_6	AS	VL	–	UV on		Y	Y	Y
A_9	–	VL	H <sub>2</sub> O <sub>2</sub>	UV on	Y			
B_7	–	VL	–	UV on	Y			
C_1	AS	–	–	UV on		Y	Y	Y
C_2	AS	–	H <sub>2</sub> O <sub>2</sub>	UV on		Y	Y	Y
C_3	AS	VL	–	–		Y		
C_4	AS	VL	H <sub>2</sub> O <sub>2</sub>	–		Y		

<sup>a</sup>: Ammonium sulfate (0.1 mM); <sup>b</sup>: vanillin (0.1 mM); <sup>c</sup>: hydrogen peroxide (11.8 mM); <sup>d</sup>: UV lamp (254 nm); <sup>e</sup>: samples collected for off-line analyses (A\_1 to A\_3 and B\_1 to B\_3 for UPLC-DAD analysis; A\_9 and B\_7 for UPLC-ToF-MS analysis); <sup>f</sup>: high-resolution time-of-flight aerosol mass spectrometric measurements; <sup>g</sup>: hygroscopic tandem differential mobility analyzer measurements; <sup>h</sup>: cloud condensation nuclei measurements. “–” indicates certain species (AS, VL, or H<sub>2</sub>O<sub>2</sub>) or element (UV) were not included in the experiments. “Y” indicates sample collection and on-line analyses using the specific instrument.

## 8. Figure Caption

Figure 1. Ratios of organics to sulfate, as well as mass fractions of organic and ammonium sulfate during experiments under conditions (A) and (B). Small grey symbols show data from all experiments (A\_1 to A\_8 or B\_1 to B\_6), while large color symbols show 10-minute running averages with error bars representing standard deviations.

Figure 2. Mass spectra at different reaction times under conditions (A) and (B), from experiments A\_2 and B\_2, respectively. The H:C and O:C ratios are also shown for each mass spectrum. Sticks with different colors indicate different ion “families”:  $C_xH_y$  being ions with only C and H,  $C_xH_yO$  being ions with C, H and only one O atom,  $C_xH_yO_z$  being ions with C, H and more than one O atoms, and HO being ions with only H and O.

Figure 3. H:C and O:C ratios as a function of time under conditions (A) and (B). Also shown are the 10-minute running average ratios of organics to sulfate (solid lines) under these two conditions.

Figure 4. The ratio of vanillin (VL) to sulfate and the ratio of vanillic acid (VA) to sulfate (panels A and B), their absolute concentrations in the aqueous solution and the natural log of their absolute concentrations (panels C and D) under the two reaction conditions. Decay rate constants ( $k_{obs}$ ) were determined at the first 50-60 minutes of the reactions. See main text for the discussion on the rate constant for the bimolecular reaction between VL and OH ( $k_{OH}$ )

for condition (A) in panel C.

Figure 5. Mass spectra of some of the products identified by UPLC-ESI-ToF-MS under condition (B).

Mass spectra of other products identified by UPLC-ESI-ToF-MS can be found in Figure S12 in the Supporting Information.

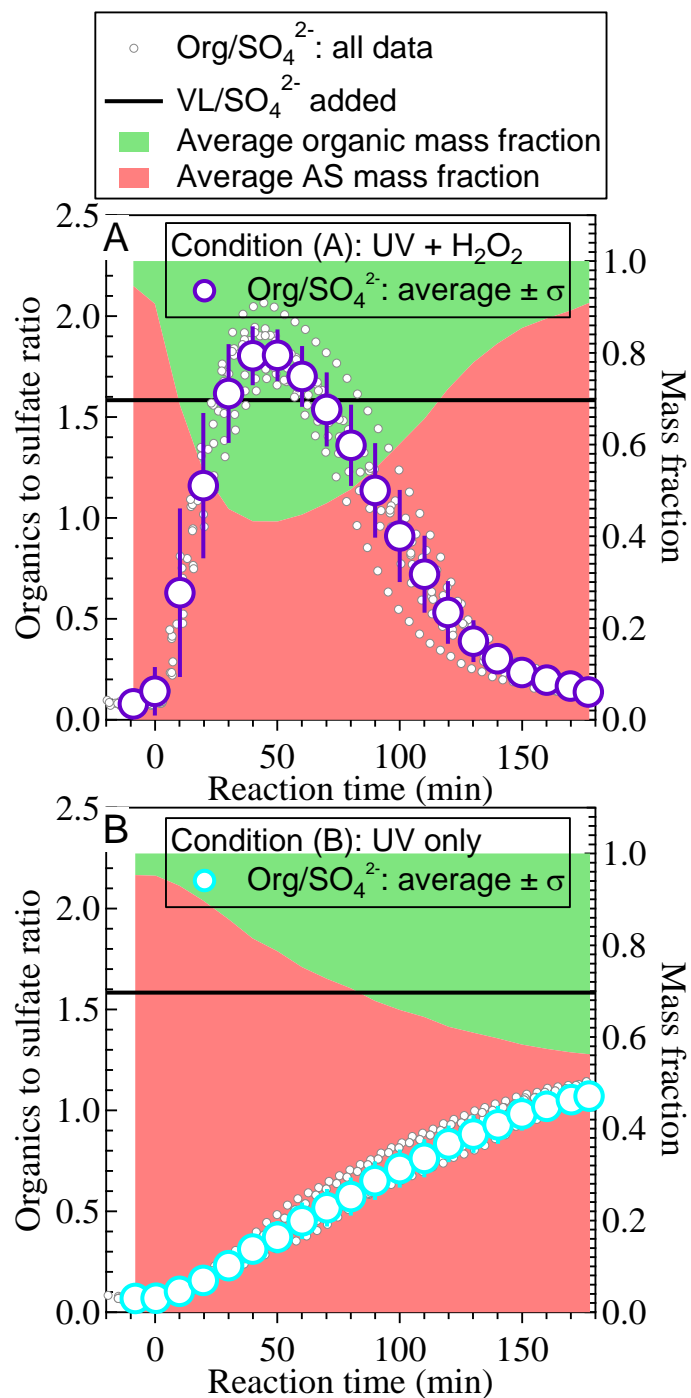
Figure 6. Mass spectra of derivatized standard compounds (shown in red) and mass spectra of derivatized samples (shown in blue) that contained small oxygenates obtained by GC-MS.

Mass spectra of other low-molecular-weight products can be found in Figure S13 in the Supporting Information.

Figure 7. HTDMA measured growth factor at 90% RH ( $GF_{90}$ ) (panels A and C) and CCNc-measured  $D_{50}$  at super-saturation (ss) of 0.1% (panels B and D) under condition (A) (panels A and B) and condition (B) (panels C and D). Also shown are volume fractions of organics and ammonium sulfate, as well as O:C ratios.

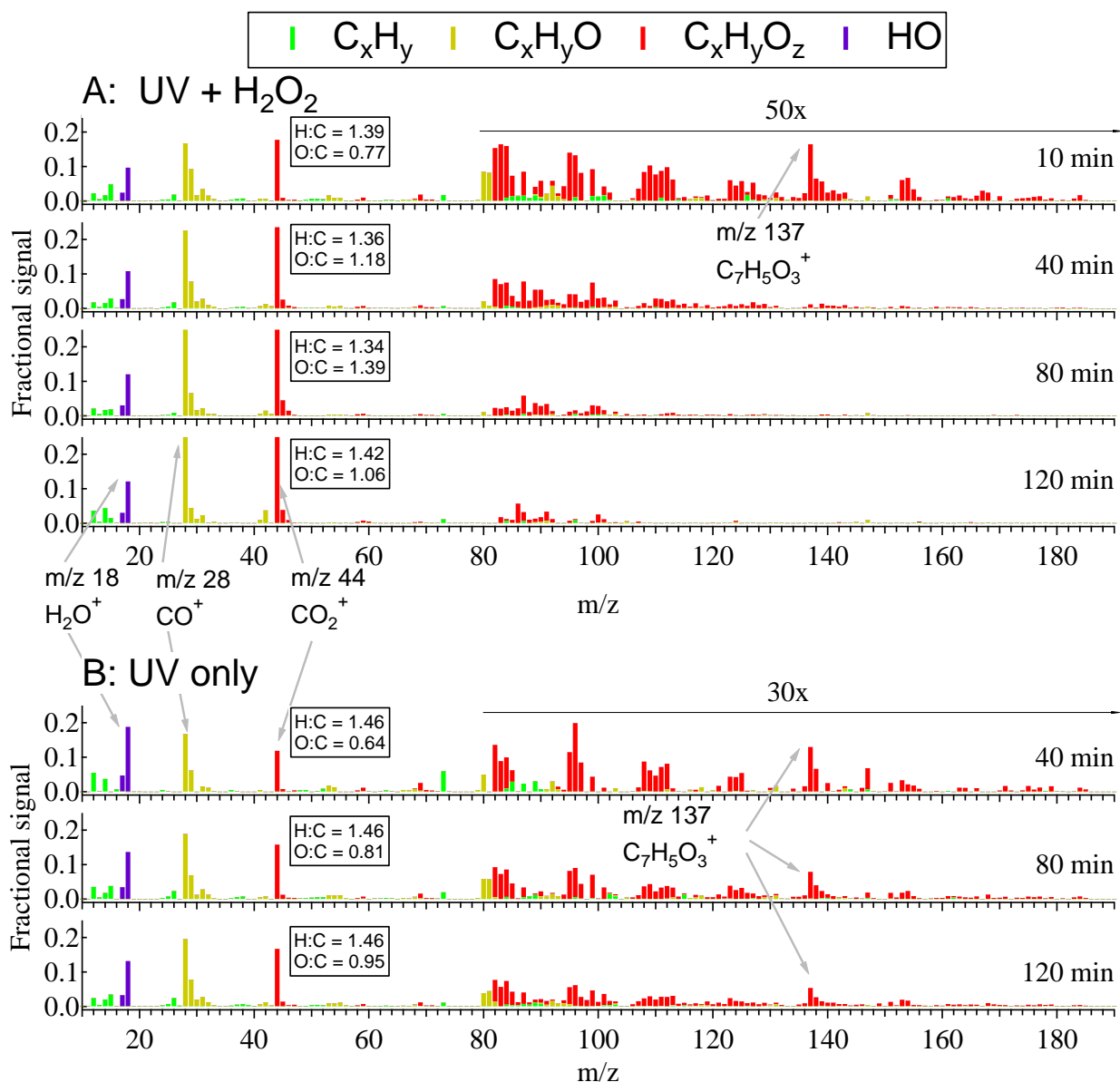
Figure 8. Organic growth factor at 90% RH ( $GF_{org}$ ) retrieved (see Supporting Information Section 6) from measured  $GF_{90}$  as a function of O:C under condition (A) (green circles) and condition (B) (blue circles). Also shown is the dependence of  $GF_{org}$  at 90% RH from Massoli et al. (2010) (black line and shaded area, which represents the uncertainty of fitting).

## 1 9. Figures



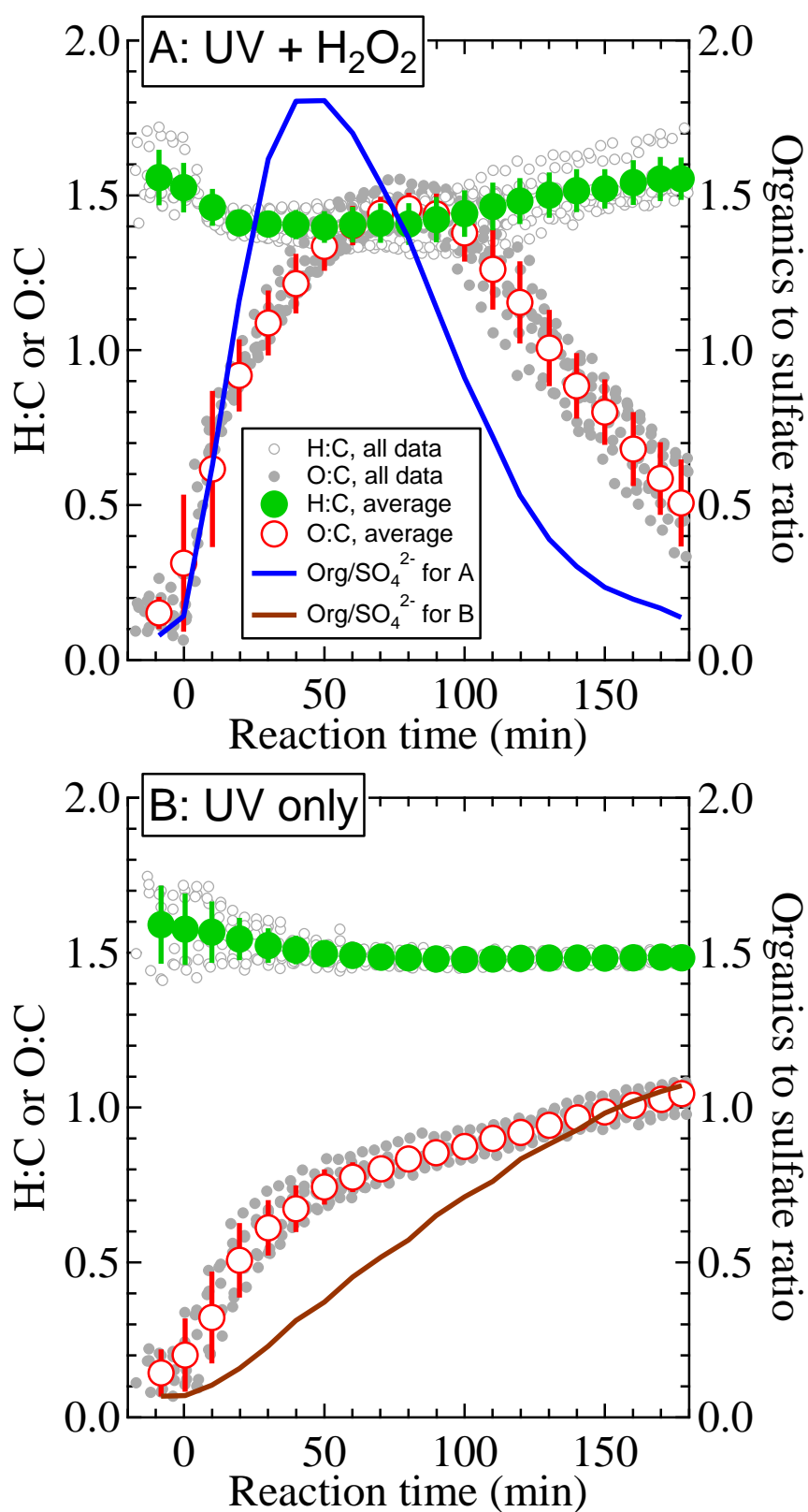
2

3 Figure 1. Ratios of organics to sulfate, as well as mass fractions of organic and ammonium sulfate  
 4 during experiments under conditions (A) and (B). Small grey symbols show data from all experiments  
 5 (A\_1 to A\_8 or B\_1 to B\_6), while large color symbols show 10-minute running averages with error  
 6 bars representing standard deviations.



1

2 Figure 2. Mass spectra at different reaction times under conditions (A) and (B), from experiments A\_2  
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 4 different colors indicate different ion “families”:  $C_xH_y$  being ions with only C and H,  $C_xH_yO$  being  
 5 ions with C, H and only one O atom,  $C_xH_yO_z$  being ions with C, H and more than one O atoms, and  
 6  $HO$  being ions with only H and O.



1

2 Figure 3. H:C and O:C ratios as a function of time under conditions (A) and (B). Also shown are the  
 3 10-minute running average ratios of organics to sulfate (solid lines) under these two conditions.

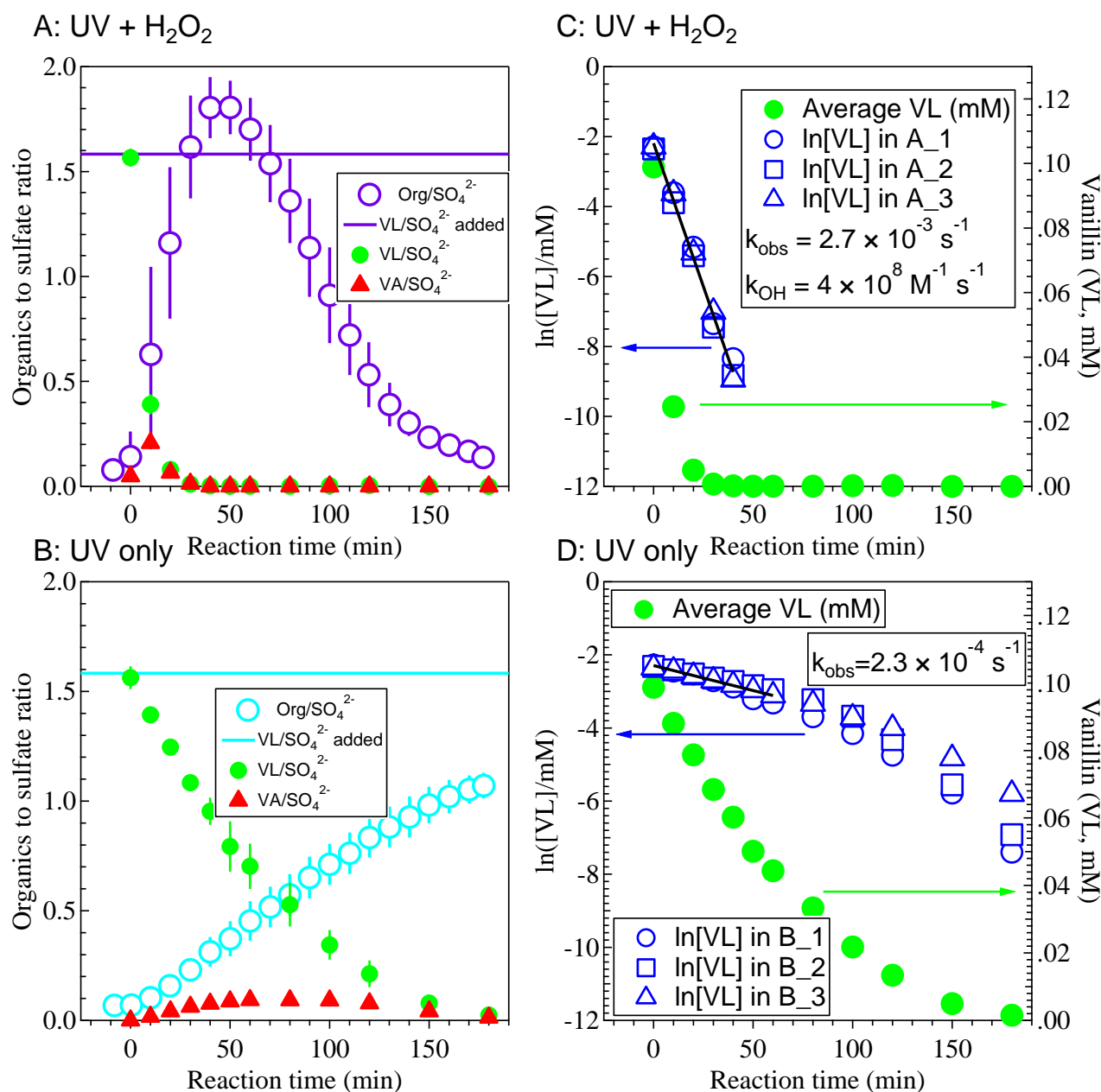
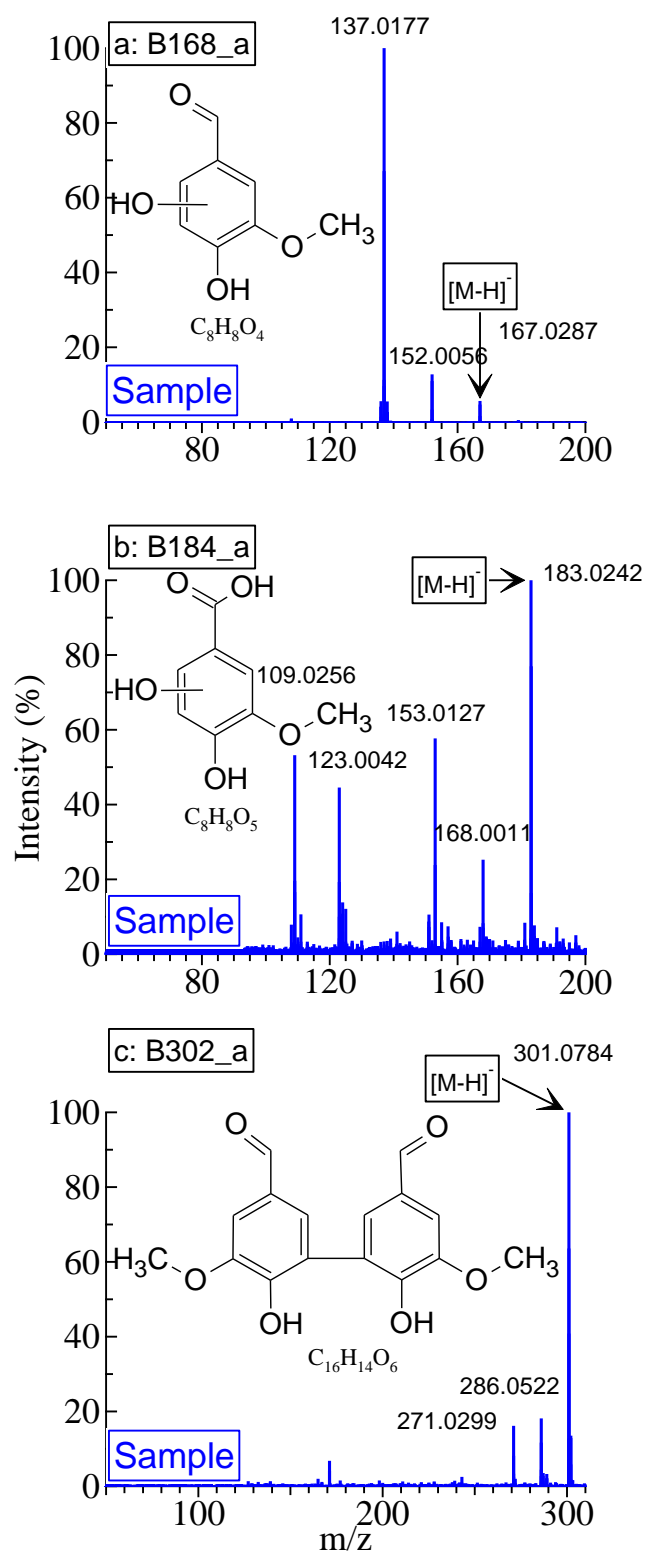


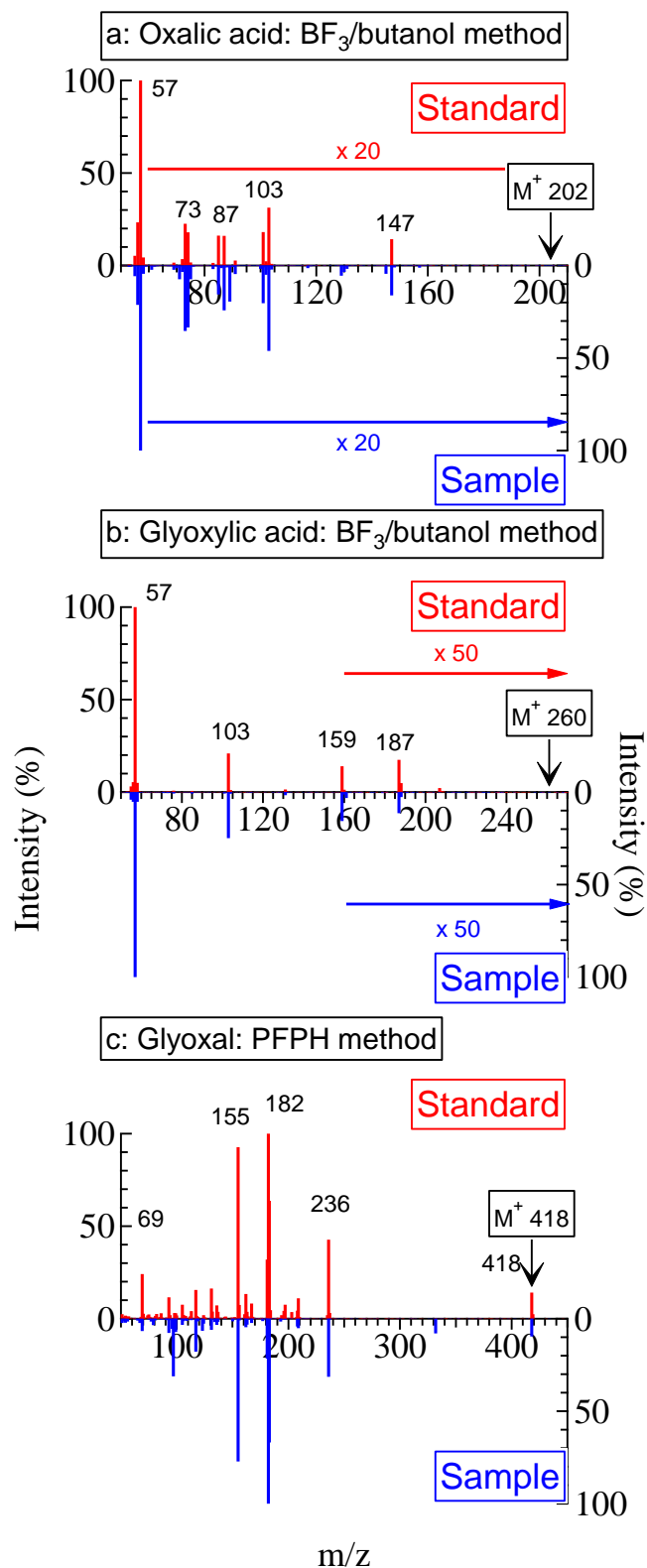
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1

2 Figure 5. Mass spectra of selected products identified by UPLC-ESI-ToF-MS under condition (B).  
 3 Mass spectra of other products by UPLC-ESI-ToF-MS can be found in Figure S12 in the Supporting  
 4 Information.



1

2 Figure 6. Mass spectra of derivatized standard compounds (shown in red) and mass spectra of  
 3 derivatized samples (shown in blue) that contained those low-molecular-weight products by GC-MS.  
 4 Mass spectra of other low-molecular-weight products can be found in Figure S13 in the Supporting  
 5 Information.

6

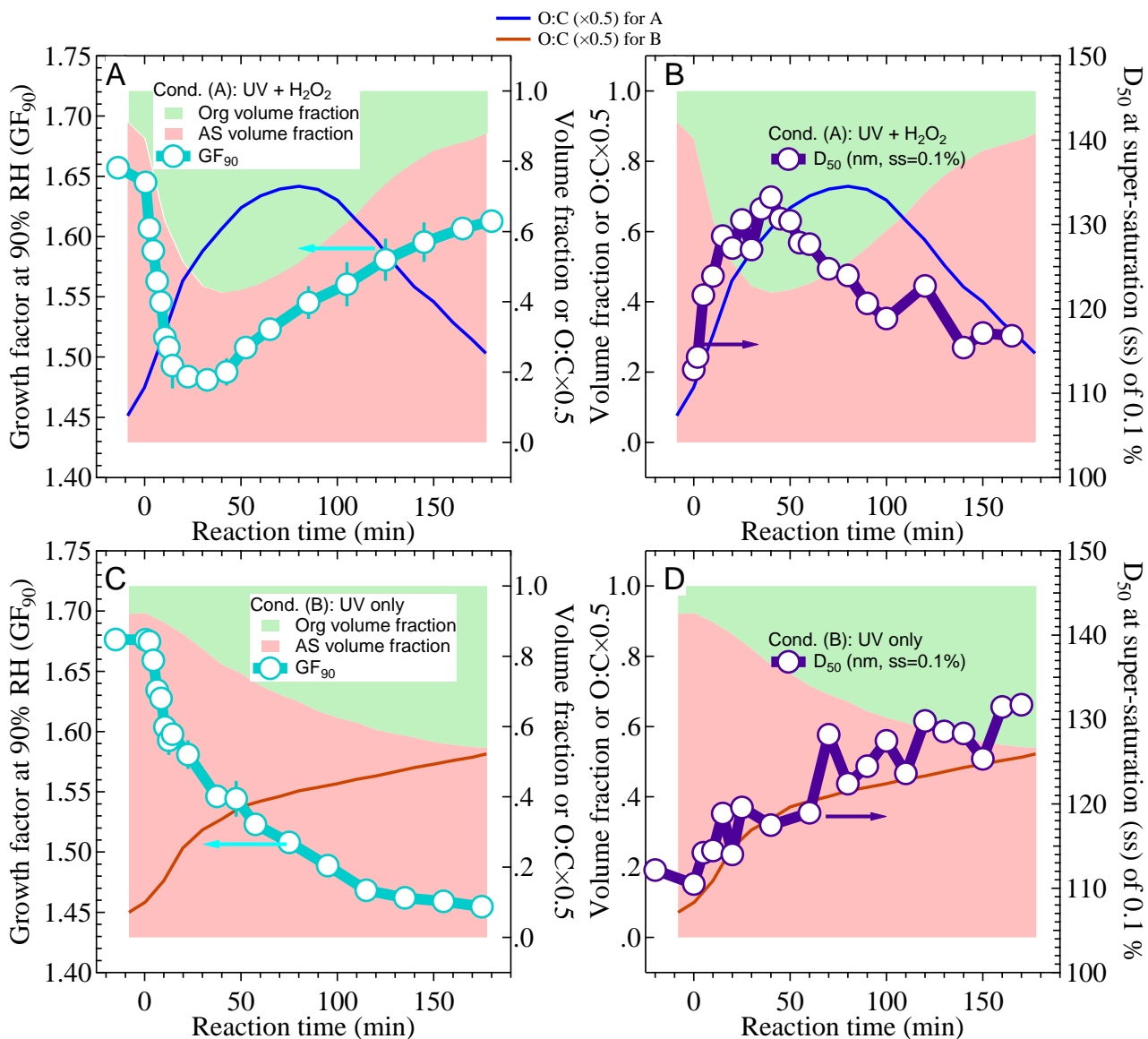
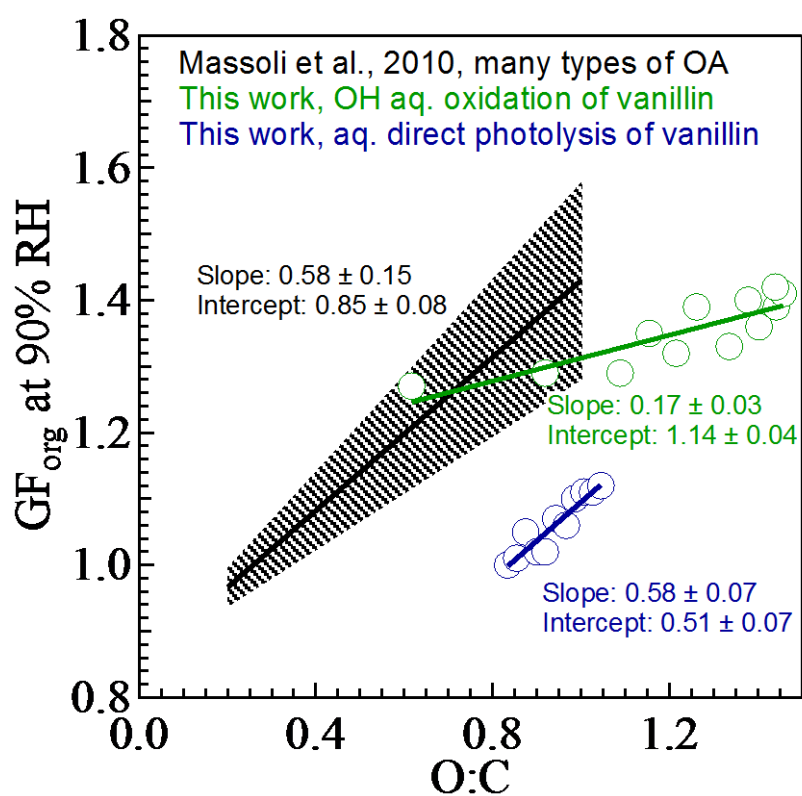


Figure 7. HTDMA measured growth factor at 90% RH ( $GF_{90}$ ) (panels A and C) and CCNc measured  $D_{50}$  at super-saturation (ss) of 0.1% (panels B and D) under condition (A) (panels A and B) and condition (B) (panels C and D). Also shown are volume fractions of organics and ammonium sulfate, as well as O:C ratios.



1

2 Figure 8. Organic growth factor at 90% RH (GF<sub>org</sub>) retrieved (see Supporting Information Section 6)  
 3 from measured GF<sub>90</sub> as a function of O:C under condition (A) (green circles) and condition (B) (blue  
 4 circles). Also shown is the dependence of GF<sub>org</sub> at 90% RH from Massoli et al. (2010) (black line and  
 5 shaded area, which represents the uncertainty of fitting).