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Dynamic light absorption of biomass burning organic carbon photochemically aged under natural sunlight

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Abstract

Wood burning aerosol produced under smoldering conditions was photochemically aged with different relative humidity (RH) and NO_x conditions using a 104 m^3 dual outdoor chamber under natural sunlight. Light absorption of organic carbon (OC) was measured over the course of photooxidation using a UV–visible spectrometer connected to an integrating sphere. At high RH, the color decayed rapidly. NO_x slightly prolonged the color of wood smoke, suggesting that NO_x promotes the formation of chromophores via secondary processes. Overall, the mass absorption cross-section (integrated between 280 nm and 600 nm) of OC increased by 11–54 % (except high RH) in the morning and then gradually decreased by 19–68 % in the afternoon. This dynamic change in light absorption of wood burning OC can be explained by two mechanisms: chromophore formation and sunlight bleaching. To investigate the effect of chemical transformation on light absorption, wood smoke particles were characterized using various spectrometers. The intensity of fluorescence, which is mainly related to polycyclic aromatic hydrocarbons (PAHs), rapidly decreased with time indicating the potential bleaching of PAHs. A decline of levoglucosan concentrations evinced the change of POA with time. The aerosol water content measured by Fourier transform infrared spectroscopy showed that wood burning aerosol became less hygroscopic as photooxidation proceeded. A similar trend in light absorption changes has been observed in ambient smoke aerosol originating from the 2012 County Line Wildfire in Florida. We conclude that the biomass burning OC becomes less light absorbing after 8–9 h sunlight exposure compared to fresh wood burning OC.

1 Introduction

Atmospheric aerosols that can both absorb and scatter sunlight remain the largest source of uncertainty in radiative forcing of climate (Forster et al., 2007). Among various atmospheric aerosols, only carbonaceous aerosol and some mineral dusts, have the

potential to cause climate warming by directly absorbing solar radiation. Carbonaceous aerosol includes both organic carbon (OC) and blank carbon (BC) aerosol. In this study, BC and elemental carbon (EC) are treated as equivalent. BC is well known for its climate warming effect. Light absorption of OC has been poorly understood mainly

5 due to its complex chemical composition and dynamic evolution under atmospheric conditions. OC can be either primary (POA) or secondary (SOA) in origin. On a global scale, approximately 69 % of POA and 23 % of SOA are contributed by biomass burning (Hallquist et al., 2009). Considering the large source of biomass burning OC in the atmosphere, better understanding its light absorption properties is necessary.

10 Recently, there is growing evidence that biomass burning OC may be a nonnegligible contributor to light absorption in atmospheric aerosols, especially in the short wavelength visible and ultraviolet spectral regions (Hoffer et al., 2006; Kirchstetter and Thatcher, 2012; Lack et al., 2012). Kirchstetter and Thatcher (2012) analyzed the filter-based wood smoke aerosol from night-time ambient sampling, and estimated that OC

15 accounted for 49 % of total aerosol light absorption between 300 nm and 400 nm. Lack et al. (2012) measured light absorption of biomass burning aerosol originating from a wildfire event using a photo-acoustic aerosol absorption spectrometer. Their results suggested that the primary biomass burning OC contributed 27(±15) % of the total absorption at 404 nm.

20 Light absorption of OC is mainly determined by its chemical composition. Biomass burning OC has a yellowish to brown color, with absorption mainly at UV range, and decreasing absorption toward visible range (Kirchstetter et al., 2004; Schnaiter et al., 2005). Chen and Bond (2010) measured light absorption of solvent-extracted fresh primary OC generated in a nitrogen filled furnace. They proposed that chemicals with

25 large, polar, and conjugated aromatic rings would be the light absorbing materials. The study of Del Vecchio and Blough (2004) suggested that hydroxy-aromatic compounds and quinoid might be responsible for OC color.

In the atmosphere, the chemical composition of OC aerosol changes with aging through processes such as oxidation of organic gases, heterogeneous oxidation with

atmospheric oxidants (e.g., ozone and OH radicals), and aerosol phase reactions. Recent laboratory experiments indicate that photochemical oxidation results in the dynamic evolution of chemical and physical properties of OC. The aged OC aerosol became less volatile and more oxygenated after a few hours of photochemical oxidation (Grieshop et al., 2009). Aging processes also change the hygroscopicity. The hygroscopic water content in OC aerosol may increase or decrease depending on wood burning conditions (Martin et al., 2013).

The aging effects on the light absorbing properties of OC have been varied among research reports due to the different aging times and oxidation conditions. For example, Adler et al. (2011) derived the effective broadband refractive index of biomass burning aerosol from a wood burning event using a white light optical counter. They found that the aged biomass burning aerosols were less absorbing than the freshly emitted aerosols, with the imaginary refractive index decreasing from 0.04 to 0.02. However, Saleh et al. (2013) recently reported that aged OC is more absorbing than fresh one due to SOA formation. In their study, the wood smoke aerosol was aged only for one hour using an indoor chamber. These conflicting results show that the influence of aging on light absorption of wood smoke OC needs further investigation.

The objectives of this study are: (1) to investigate the influence of photochemical oxidation on light absorption of wood smoke OC; (2) to explore the effects of relative humidity (RH) and NO_x on light absorption of wood smoke OC over the course of the photochemical aging process, and (3) to characterize the chemical evolution of OC aerosol. In order to focus on wood burning OC with minimal influence by BC, wood smoke was produced from smoldering-phase burning. Smoldering combustion has frequently been found in biomass burning situations including wildfires, prescribed burns, and agriculture burning (Alves et al., 2011; Hille and Stephens, 2005; Hays et al., 2005). It has been reported to consume over 50 % of biomass in temperate and boreal fires (Bertschi et al., 2003). In this study, the resulting wood burning OC was photochemically oxidized using a large outdoor chamber (104 m^3) that exposure

to natural sunlight. The large outdoor chamber allowed atmospheric aging conditions to be closely mimicked and aerosol samples to be collected for longer time periods.

2   Methods

2.1   Outdoor chamber experimental setup

The photooxidation of wood smoke was performed using the University of Florida Atmospheric Photochemical Outdoor Reactor (UF-APHOR) dual chambers ($52\,m^3$ per chamber) which are located on the roof of Black Hall, UF. The UF-APHOR dual chambers with half-cylinder design are made of FEP Teflon film attached to metal frames. A detailed description of the UF-APHOR chambers can be found elsewhere (Im et al., 2013).

Prior to each experiment, the chambers were continuously flushed with clean ambient air for more than 12 h and then purged overnight with clean air from an air cleaner (GC Series, IQAir Inc.). Commercial hickory hardwood logs were chopped into pieces with approximate size of $3\,cm \times 4\,cm \times 33\,cm$. The wood pieces were burned under smoldering conditions using a wood grill stove. Before sunrise, smoke was introduced into the chambers through a 4 inches alumina tube (0.5 m in length) connected to the stove. Immediately after injecting wood smoke, the chamber air was mixed for five minutes using mixing fans. The initial conditions for wood smoke particle concentration measured by an organic carbon (OC) analyzer are summarized in Table 1. For high NO_x experiments, NO gas from the tank was introduced into the chamber after the wood smoke injection.

Photochemical reactions started at sunrise. The sunlight intensity, chamber temperature, and chamber relative humidity (RH) dynamically changed over the course of chamber experiments. A typical time profile of these parameters is shown in Fig. S1. To study of the effect of humidity on photochemical aging of wood smoke, the dry condition humidity (RH < 30 %) was controlled one day prior to the experiment using a dehumid-

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ifier. To achieve the wet condition (RH > 60 %), the chamber air was humidified using a steam evaporator during the experiment. Sunlight intensity and chamber temperature, and chamber humidity were measured continuously with an ultraviolet radiometer (model TUVR, The Eppley Laboratory) and a Temp-RH sensor (CS2, Campbellscientific), respectively.

2.2 Wood smoke characterization

Gas and particles were sampled through manifolds that run directly to a laboratory just below the chambers. The wood smoke particles were collected on a 13 mm diameter filter (Borosilicate microfibers reinforced with woven glass cloth and bonded with Teflon, Gelman Science Pallflex, Type: TX40H120-WW) at a flow rate of 17 L min^{-1}. The UV–visible light absorption spectrum of aerosols collected on the filter was directly monitored using a Perkin–Elmer Lambda 35 UV–visible spectrophotometer equipped with a Labsphere RSA-PE-20 diffuse reflectance accessory. The detailed measurement procedure for light absorption spectra of aerosols can be found in our previous study (Zhong and Jang, 2011). OC and elemental carbon (EC) of the aerosol filter samples were analyzed with a semi-continuous OC/EC analyzer (Model 4, Sunset Laboratory Inc.). Levoglucosan obtained from the filter/acetonitrile extract was first derivatized with N,O-bis(trimethylsilyl) trifluoroacetamide(BSTFA) and then measured using a gas chromatograph/ion trap mass spectrometer (GC-ITMS) (CP-3800 GC, Saturn 2200 MS, Varian Inc.). The fluorescence spectra of extract of the filter sample with dichloromethane were measured by a fluorescence spectrophotometer (F-2500, Hitachi, Ltd.). The chemical functional group of particles impacted on a silicon disk was characterized using Fourier transform infrared spectroscopy (FTIR) (Nicolet Magna 560, SpectraLab Scientific Inc). To determine the water content of aerosols, the sampled silicon disk was installed in a miniature flow chamber which was located in the FTIR optical beam path. The relative humidity inside the flow chamber was controlled by passing wet air through the chamber. FTIR spectra of aerosols on the disk were obtained under varying relative humidity from 10 % to 85 %. Detailed descriptions of the

methods can be found in the previous study by Jang et al. (2010). Particle concentrations and particle size distributions were monitored continuously using a scanning mobility particle sizer (Model 3080, TSI Inc.), together with a condensation nuclei counter (Model 3025A, TSI Inc.). Ozone was monitored using a photometric ozone analyzer (model 400E, Teledyne Instruments), and NO_x concentrations were measured using a chemiluminescence NO/NO_x analyzer (model 200E, Teledyne Instruments).

2.3 Light absorption of ambient aerosol

In April 2012, the County Line Fire located in Pinhook Swamp, Florida, 90 miles north of UF, burned more than ten thousand acres. Biomass burning aerosol was collected at the UF sampling site on 9–10 April 2012. During the sampling day, the OC concentration in the morning was as high as $92 \, \mu g \, m^{-3}$ (measured at the sampling site) compared to a normal day in which the OC value is 2 to $4 \, \mu g \, m^{-3}$. Thus, wildfire emission is the main source of the ambient aerosols in UF area during this episode. Fine particulate matter ($PM_{2.5}$) passing through a cyclone were collected on a 13 mm diameter filter at a flow rate of $16.7 \, L \, min^{-1}$. Light absorption spectra of the filter sample were directly measured using a UV–visible spectrometer equipped with an integrating sphere.

3 Results and discussion

3.1 Light absorption of OC

A wood burning particle contains a mixture of BC and OC. Due to its strong light absorption capacity, a small amount of BC may contribute significant spectral absorption depending upon wavelength. Hence, burning conditions were controlled as smoldering to generate less BC. The percentage of EC in smoldering smoke particles was near or less than 5 % as shown in Table 1. A method introduced by Kirchstetter and Thatcher (2012) was used to separate light absorption by OC from the absorption of

the smoke aerosol mixture (OC and BC). Light absorption of OC, Abs_{OC}, is obtained by subtracting the absorption of BC, Abs_{BC}, from the mixture's absorption Abs:

$$Abs_{OC} = Abs - Abs_{BC} \tag{1}$$

Abs is directly measured from the filter sample. Abs_{BC} is estimated as:

$$Abs_{BC} = \frac{Abs(880\,nm)}{880\,nm^{-AAE}}\lambda^{-AAE} \tag{2}$$

where AAE is the Absorption Angstrom Exponent and defined as:

$$AAE = \frac{-\ln(Abs_{BC}(\lambda_1)/Abs_{BC}(\lambda_2))}{\ln(\lambda_1/\lambda_2)} \tag{3}$$

For Abs_{BC} estimation, the measured AAE value was 1.03 for wavelengths between 360 and 900 nm.

Figure 1a shows absorption spectra of the fresh smoke aerosol, the decoupled OC, and the decoupled BC. Over the wavelength range between 280 nm and 600 nm, OC contributes 60–98 % of the total sample absorption, while BC is dominate for longer wavelengths (> 700 nm). Light absorption of OC exhibits no distinct peaks, exponentially decreasing with increasing wavelength. In the visible range (400–700 nm), the average AAE of OC is 4.74. This value accords with the AAE of ambient biomass burning OC (an average value of 5) reported by Kirchstetter and Thatcher (2012).

All samples were analyzed in a similar manner as described above. The main data reported here are mass absorption cross-section of OC (MAC_{OC}, $m^2\,g^{-1}$), which is wavelength (λ) dependent and determined by normalizing the absorption coefficient of OC ($b_{abs(OC)}$, m^{-1}) with thermal OC mass concentration, M_{OC} ($g\,m^{-3}$, measured by an OC/EC analyzer)

$$MAC_{OC} = \frac{b_{abs(OC)}}{M_{OC}} \tag{4}$$

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In Eq. (4), $b_{abs(OC)}$ can be calculated from Abs_{OC}:

$$b_{abs(OC)} = \frac{1}{f}\frac{A}{V} Abs_{OC} \ln(10) \qquad (5)$$

where f is the correction factor with a value of 1.4845 obtained from our previous study (Zhong and Jang, 2011) for the purpose of eliminating the absorbance caused by filter material scattering, A is the filter surface area sampled ($7.85 \times 10^{-5}\,m^2$ in this study), and V is the volume (m^3) of air passing through the filter during a given sampling time. Abs_{OC} is obtained from Eq. (1).

3.2 Effect of photochemical aging on light absorption of OC

Our chamber experimental data showed that light absorption of wood burning OC was significantly modified due to photochemical aging. Figure 1b shows the typical MAC_{OC} of wood smoke exposed to sunlight using the outdoor chamber at different daytime hours. Overall, MAC_{OC} increased with aging time in the morning. The total MAC_{OC}, which is estimated as the area under the MAC_{OC} spectrum curve between 280 nm and 600 nm, increased by 11–54 % (26 % on average) except data from the high RH condition (West chamber on 24 October 2012). Saleh et al. (2013) have explained that the increment in aged OC absorption is caused by the formation of SOA, which absorbs more than fresh POA in the short wavelength visible and near-UV regions. Schauer et al. (2001) reported that about 34 % of identified organic compounds in the gas phase and 41 % of organic compounds identified in POA were phenols, syringols, and guaiacols. These phenols and methoxylated phenols in the gas phase or in the particle can rapidly react with atmospheric oxidants to form light brown substances (Gelencser et al., 2003; Chang and Thompson, 2010; Ofner et al., 2011). Thus, the enhancement in light absorption would be caused by either oxidized POA or SOA produced from oxidation of primary phenolic compounds.

However, light absorption of OC at both UV and visible ranges gradually decreased in the afternoon. The total MAC_{OC} decreased by 19–68 % (41 % on average) at the end

of the experiment. The possible explanation for the decreases in MAC_{OC} is bleaching of colorant by sunlight. The previous study by Zhong and Jang (2011) reported the bleaching effect of light on light absorption of SOA. In wood smoke POA and SOA, there are abundant of chromophores, such as conjugated aromatic rings and phenols, as well as nitro and hydroxyl groups. High energy photons in sunlight can excite electrons in colored molecules through $\pi - \pi^*$ or $n - \pi^*$ transitions, and disrupt the conjugated structure of chromophores, resulting in the gradual fading of wood smoke color.

3.3 Effect of RH on light absorption of OC

In order to investigate the effect of humidity on MAC_{OC}, two sets of wet-dry dual chamber experiments were conducted: high RH (80–87 %) vs. low RH (10–26 %) on 24 October 2012, and middle RH (56–75 %) vs. low RH (12–23 %) on 1 November 2012. Figure 2 illustrates the time profile of the relative MAC_{OC} and MAC_{OC} at 550 nm. The relative MAC_{OC} was obtained by normalizing the total MAC_{OC} at a certain time by the initial total MAC_{OC}. As photooxidation progressed in the afternoon, the absorption decay at wet conditions became more rapid than the decay at dry conditions. For example, at the end of the experiment, the relative MAC_{OC} and MAC_{OC} at 550 nm at the low RH were higher by 43 % and 45 %, respectively, compared to those at the high RH. Unlike MAC_{OC} values in the middle or the low RHs, no increase appeared for MAC_{OC} values in the high RH experiments. Elevated values of relative humidity can increase the water content of wood smoke particles (see Sect. 3.5.4). Anastasio et al. (1997) reported that illumination of aqueous phase non-phenolic aromatic carbonyls in the presence of phenols can produce H_2O_2 and destroy phenols. Since both the aromatic carbonyls and phenols are major products from the combustion of wood (Rogge et al., 1998), this aqueous phase photooxidation would produce a significant amount of H_2O_2, In wet aerosols, H_2O_2 can photodisssociate to produce OH radicals (Zellner et al., 1990; Faust, 1994), which could decompose chromophores and lead to bleaching of wood OC.

3.4 Effect of NO$_x$ on light absorption of OC

Controlled dual chamber experiments were also conducted to study the influence of NO$_x$ on MAC$_{OC}$ of wood aerosols, high NO$_x$ (108 ppb) vs. low NO$_x$ (16 ppb) on 30 October 2012 and middle NO$_x$ (43 ppb) vs. low NO$_x$ (16 ppb) on 11 October 2012. Figure 3 shows the time profiles of the relative MAC$_{OC}$ values and MAC$_{OC}$ at 550 nm under different NO$_x$ concentrations. In the presence of high NO$_x$, the MAC$_{OC}$ values are slightly higher than those at low NO$_x$ concentration. NO$_x$ is able to modify the reaction pathway for organic compounds in the gas phase. As discussed in Sect. 3.1 (light absorption of OC), nitro-phenols (e.g., nitrocatechols) and nitro compounds can be produced through the photooxidation of phenolic organic compounds and contribute SOA. Nitrophenols have been suggested to be strong candidates to represent light absorbing toluene SOA (Nakayama et al., 2010; Zhong et al., 2012; Zhang et al., 2013). At higher NO$_x$ level, the concentration of these chromophoric nitro-phenols can be greater, decelerating the decay of MAC$_{OC}$ of wood smoke OC.

3.5 Chemical and physical evolution of organic carbon aerosol

To investigate how the chemical composition of wood smoke aerosol influences light absorption, the aerosol introduced into the chamber was characterized as photooxidation progressed. In this study, levoglucosane, a major constituent in wood smoke POA, was monitored by GC-ITMS. The intensity of fluorescence, which is mainly due to polycyclic aromatic hydrocarbons (PAHs), was measured for solvent extracted aerosol samples. To study the alteration of the hygroscopic properties of wood smoke aerosol due to photooxidation, the aerosol water content was also measured by FTIR.

3.5.1 Levoglucosan decay

Levoglucosan is abundant in wood smoke aerosol, contributing 3–49 % of total wood smoke aerosol carbon (Mazzoleni et al., 2007). Figure 4a illustrates the rapid degrada-

tion of levoglucosan in both pure levoglucosan aerosol and wood smoke aerosol after wall–loss correction of aerosols. The experiment method for the pure levoglucosan experiment is provided in the Supplement (document S1). The decay rate of the levoglucosan associated with wood smoke aerosol is similar to that of pure levoglucosan

5 within the experimental error. The decrease of levoglucosan would be caused by wall loss of levoglucosan vapor and photochemical oxidation by OH radicals in the aerosol. According to Booth et al. (2011), the sub-cooled liquid vapor pressure of levoglucosan is 1.9×10^{-4} pa at 298 K, thus the wall loss of gas-phase levoglucosan is likely. In order to confirm the photochemical oxidation of levoglucosan, gas and aerosol samples were

10 analyzed using GC-ITMS to identify the new products. Five products have been tentatively identified using the mass fragmentation spectra as shown in Figs. S2 and S3. The reaction pathways for the formation of the proposed products are illustrated in Fig. S4. However, the low amount of O_3 formation suggests that oxidation of levoglucosan occurred mainly via aerosol phase reaction.

15 **3.5.2 PAHs decay**

Unlike levoglucosan, the mass percentage of PAHs in wood smoke is less than 1 % (Schauer et al., 2001). PAHs absorb UV light and emit fluorescence in the visible region. The photooxidation of PAHs can therefore be examined by measuring the fluorescence as a function of time. The fluorescence spectra of OC aerosol (Fig. 4b) show

20 a rapid decrease in fluorescence emission when excited at 280 nm. Similar to levoglucosan, such decline can be caused by both evaporation combined with increased temperature after the sun rises (Fig. S1) and photochemical oxidation of PAHs. PAHs can be transformed to oxy-PAHs via aerosol phase oxidation (e.g., reaction with singlet oxygen produced by a photosensitizing process or reaction with OH radicals and free

25 radical reactions) (Jang and McDow, 1995). These oxy–PAHs generally have much lower fluorescence quantum yield compared to unoxidized PAHs.

3.5.3 FTIR spectra of wood burning aerosol

Figure 5a compares the characteristic FTIR spectra of fresh and aged wood burning particles at RH of 45 %. For the fresh particles, the O–H stretching of alcohols is seen at 3500–3200 cm^{-1}. The C=O stretching of carbonyls (e.g., aldehydes, ketones, and carboxylic acids) occurs at 1800–1680 cm^{-1}, and the aromatic C=C stretching bands appear at 1610, 1517, and 1458 cm^{-1}. After photochemical aging of wood smoke aerosol for 5 h, the absorbance for both the O–H stretch of alcohols and the aromatic C=C stretch significantly decreased. The FTIR spectrum of aged particles also showed the formation of carboxylic acids confirmed by the broad O–H stretch at 3300–2500 cm^{-1} and the increased absorbance at C=O stretching frequency, suggesting oxidation of wood constituents.

3.5.4 Hygroscopic properties of wood burning aerosol

To investigate the hygroscopicity of wood smoke particle, the water content of the particle was monitored using the FTIR equipped with a miniature flow chamber under varying RHs (Jang et al., 2010). The absorbance at 1650 cm^{-1}, which is the characteristic frequency for O–H bending of water, was used to estimate the water content in the aerosol. The peak at 1650 cm^{-1} was calibrated with the NaCl particle, which has a known water content obtained by an inorganic thermodynamic model at a given RH (e.g., AIM-III model, http://www.aim.env.uea.ac.uk/aim/aim.php) (Clegg et al., 1998). As shown in Fig. 5b, the fresh particle was much more hygroscopic than the aged particle. The water content in fresh particles exponentially increased as a function of RH. No water was detected the RH was lower 25 %. In the morning, fresh wood smoke particle contains a high amount of the sugar compounds such as levoglucosan, which is very hydrophilic and easily absorbs water. As described in Sect. 3.5.1 (Levoglucosan decay), the amount of primary sugars rapidly decayed due to photooxidation. Although the carboxylic acid content increased with aging in aged wood aerosol (FTIR data in Fig. 5a), the aerosol became less hygroscopic.

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4   Conclusion and atmospheric implication

The dynamic changes in light absorption of biomass burning OC were investigated using the outdoor smog chamber under natural sunlight. The diurnal change in absorption of wood smoke aerosol is governed by two mechanisms: chromophore formation and sunlight bleaching. The colored products originating from photooxidation of phenolic SVOCs can increase wood smoke OA absorption in the morning while sunlight fades the color of both POA and SOA in the afternoon. Higher concentrations of NO_x help prolong the wood smoke aerosol color (Fig. 3), but the high RH accelerates the degradation of aerosol color (Fig. 2). The decay of both PAHs and levoglucosan indicated dynamic changes in chemical composition of primary OC due to aging. The aged wood smoke OC became more oxidized but less hygroscopic as shown by FITR spectra measurement.

The ambient data obtained during the County Line Wildfire event on 9–10 April 2012, were analyzed using Eqs. (1)–(5) with a measured AAE of 0.74 for ambient BC. Figure 6 shows the light absorption of ambient biomass burning OC after subtracting background OC. Compared to the MAC_{OC} sampled at 08:30 a.m., the MAC_{OC} sampled at 09:36 a.m. on 9 April increased by 18 %; but the MAC_{OC} decreased by 28 % at 11:30 a.m. The diurnal pattern in light absorption of the ambient biomass burning OA is consistent with the result obtained in the outdoor chamber.

There is increased emphasis on research pertaining to the climate forcing of OC, including both experimental studies and model simulations. Considering the lifetime of aerosol (about six days), the results of both outdoor chamber studies (Fig. 1) and ambient field data (Fig. 6) suggest that biomass burning OC will absorb less light as photochemical reactions progress.

Supplementary material related to this article is available online at http://www.atmos-chem-phys-discuss.net/13/20783/2013/ acpd-13-20783-2013-supplement.pdf.

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Table 1. Summary of experimental conditions of fresh wood smoke before photochemical oxidation.

No	Date	Chamber	RH %	Temp, K	Initial NO_X, ppb	Initial OC, $\mu g\,m^{-3}$	EC / TC, %	Notes
1	11 Oct 2012	E	39–95	291–312	43	56	4.7	middle NO_X
2	11 Oct 2012	W	51–95	291–311	16	40	5.5	low NO_X
3	30 Oct 2012	E	21–87	275–300	108	58	4.1	high NO_X
4	30 Oct 2012	W	26–88	275–297	16	68	3.3	low NO_X
5	24 Oct 2012	E	10–26	291–311	5	58	3.2	low RH
6	24 Oct 2012	W	80–87	291–310	6	58	3.6	high RH
7	1 Nov 2012	E	12–23	281–306	5	111	4.3	low RH
8	1 Nov 2012	W	56–75	283–304	6	77	2.8	middle RH
9	21 Nov 2012	E	67–87	285–305	10	144	2.5	chemical analysis
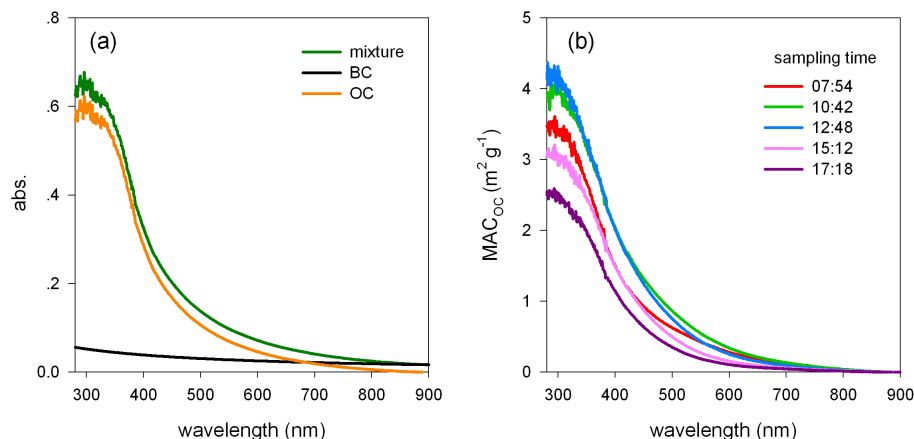
Fig. 1. (a) The decoupled absorbance spectrum of wood smoke OC obtained by subtracting the BC absorbance from the total absorbance of wood burning particles. **(b)** The diurnal pattern of MAC_{OC} of wood smoke OC photochemically aged under natural sunlight (11 October 2012, middle NO_x).
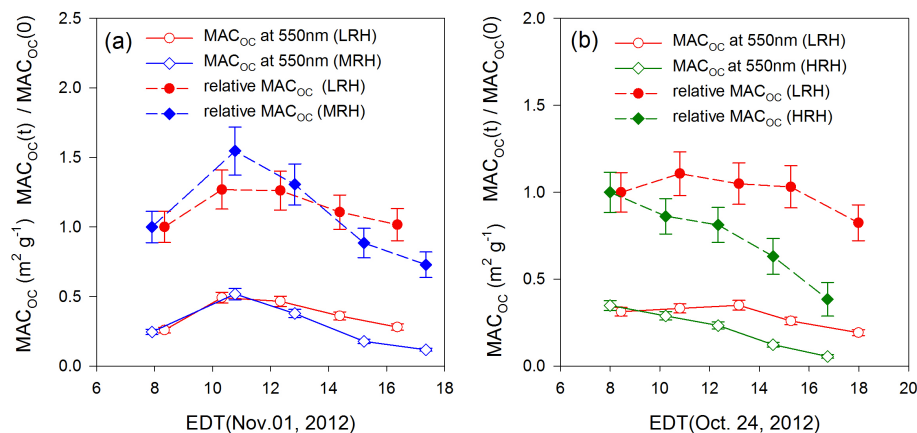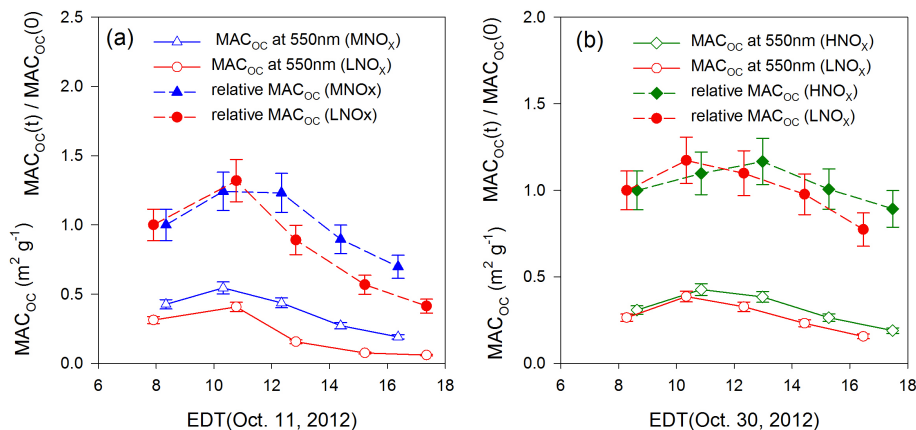
Fig. 2. Comparison of light absorption of wood OA photochemically oxidized at different conditions: **(a)** low RH vs. middle RH, **(b)** low RH vs. high RH. The open symbols are for MAC_{OC} at 550 nm. The filled symbols are for the relative MAC_{OC}, which is expressed as the total MAC_{OC} divided by the initial total MAC_{OC}. The error associated with MAC_{OC} was estimated based on the instrumental errors from OC/EC analyzer and UV–visible spectrometer as well as the uncertainty from the correction factor f, (see Eq. 5).

Fig. 3. Comparison of light absorption of wood OA photochemically oxidized at different conditions **(a)** low NO_x vs. middle NO_x, **(b)** low NO_x vs. high NO_x. The open symbols are for MAC_{OC} at 550 nm. The filled symbols are for the relative MAC_{OC}, which is expressed as the total MAC_{OC} divided by initial total MAC_{OC}. The error associated with MAC_{OC} was estimated based on the instrumental errors from OC/EC analyzer and UV–visible spectrometer as well as the uncertainty from the correction factor, f (see Eq. 5).
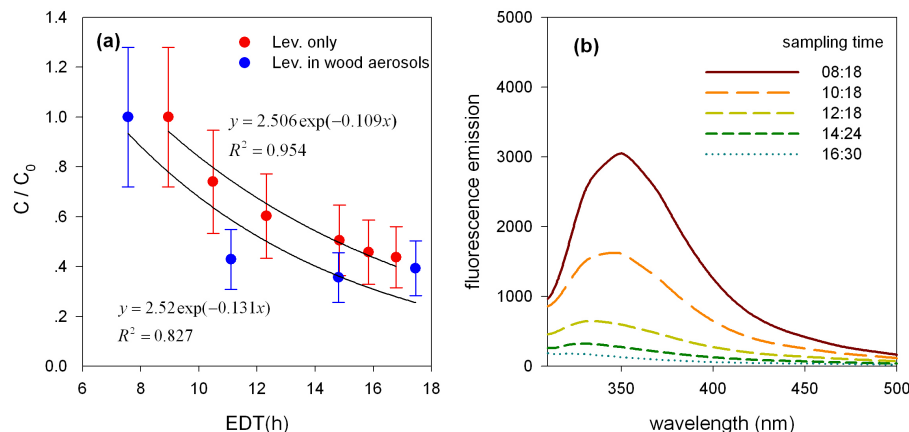
Fig. 4. Time profile for the decay of pure levoglucosan (27 March 2013) and levoglucosan in wood smoke OA (21 November 2012). The concentrations of levoglucosane (C) were corrected for wall-loss of particles and then normalized by the initial leveglucosan concentration (C_0). The associated error with the measured concentration of levoglucosan by GC–MS is ±20 %. **(b)** Fluorescence emission spectra of wood smoke particles collected at different times on 21 November 2012. The excitation wavelength was 280 nm. The error associated with the concentration of levoglucosan was estimated based on GC–MS.
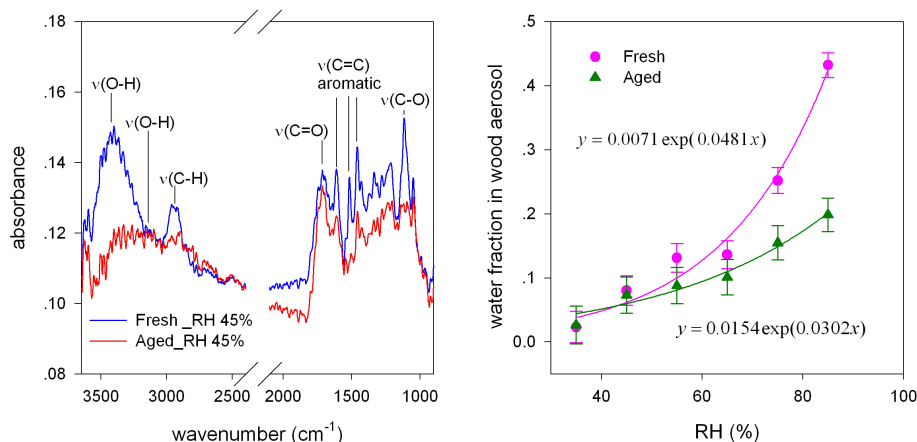
Fig. 5. (a) FTIR spectra of fresh and aged wood burning particles sampled on 21 November 2012. The spectra were recorded at RH of 45 %. **(b)** The water content of fresh particles and aged particles as a function of RH. The water content was measured with decreasing RH. The error associated with the water fraction in aerosol was estimated based on aerosol mass and FTIR absorbance at 1650 cm^{-1}.
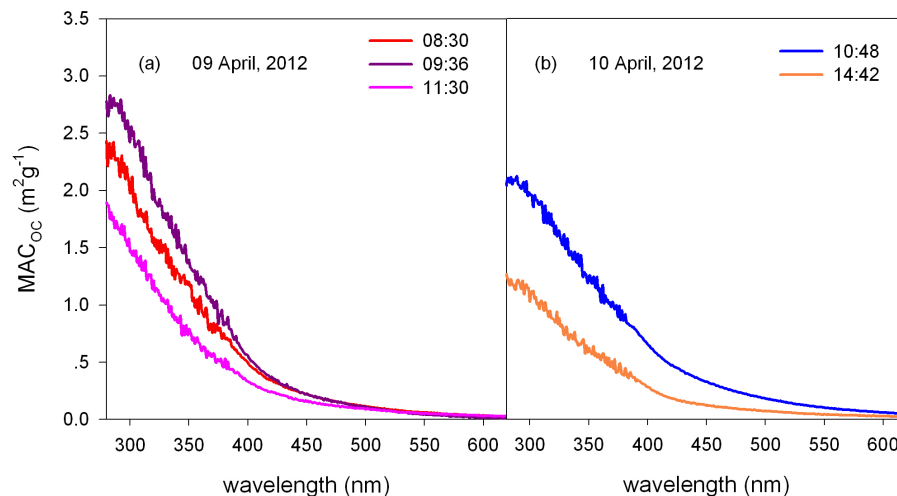
Fig. 6. Light absorption of ambient biomass burning OA sampled during the County Line Wildfire event on 9 April 2012 **(a)** and 10 April 2012 **(b)** at different sampling times.

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