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Amine reactivity with charged sulfuric acid clusters

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Abstract

The distribution of ionic species produced by electrospray of an ammonium sulfate solution in both positive and negative polarities is examined using Fourier transform ion cyclotron resonance mass spectrometry (FTICR-MS). Positively-charged ammonium bisulfate cluster composition differs significantly from negatively-charged cluster composition. For positively-charged clusters all sulfuric acid is neutralized to bisulfate, whereas for negatively-charged clusters the degree of sulfuric acid neutralization is cluster size-dependent. With increasing cluster size (and, therefore, a decreasing role of charge), both positively- and negatively-charged cluster compositions converge toward ammonium bisulfate. The reactivity of negatively-charged sulfuric acid-ammonia clusters with dimethylamine and ammonia are also investigated by FTICR-MS. Two series of negatively-charged clusters are investigated: $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_x]^-$ and $[(\text{NH}_4)_x(\text{HSO}_4)_{x+1}(\text{H}_2\text{SO}_4)_3]^-$. Dimethylamine substitution for ammonia in $[(\text{NH}_4)_x(\text{HSO}_4)_{x+1}(\text{H}_2\text{SO}_4)_3]^-$ clusters is nearly collision-limited, and subsequent addition of dimethylamine to neutralize H_2SO_4 is within one order of magnitude of the substitution rate. Dimethylamine addition to $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_x]^-$ clusters is either not observed or very slow. The results of this study indicate that amine chemistry will be evident and important only in large ambient negative ions ($> m/z 400$), whereas amine chemistry may be evident in small ambient positive ions. Addition of ammonia to unneutralized clusters occurs at a rate that is $\sim 2\text{--}3$ orders of magnitude slower than incorporation of dimethylamine either by substitution or addition. Therefore, in locations where amine levels are within a few orders of magnitude of ammonia levels, amine chemistry may compete favorably with ammonia chemistry.

1 Introduction

The process of forming small clusters from gas-phase molecules and their subsequent growth to a size where they may serve as cloud condensation nuclei (CCN) and thereby

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affect global climate is poorly understood. One reason for this uncertainty is that particle formation occurs in the 1 nm diameter size range, whereas aerosols are activated to serve as CCN around 100 nm diameter. Growth from the nucleating cluster to the CCN size range therefore requires a 10^6 -fold increase in particle mass. This process may be a significant contributor to global CCN levels (Merikanto et al., 2009). Enhancing our ability to predict aerosol climate effects is critical for a better quantitative understanding of the contributors to radiative forcing (Solomon et al., 2007). A better understanding of aerosol climate effects requires improved capability in prediction of ambient CCN levels under varying conditions (Kuang et al., 2010; Spracklen et al., 2008). Such predictive capability necessitates a mechanistic understanding of new particle formation.

The early steps of particle formation and growth have recently attracted much scientific interest (Bzdek and Johnston, 2010). In particular, amines have emerged as potentially significant contributors to these processes. This interest is due in part to the observation of amines during the growth of newly nucleated particles (Smith et al., 2010, 2008; Makela et al., 2001). Additionally, computational studies predict particulate amine levels may be enhanced relative to ammonia, despite the much lower atmospheric concentrations of amines (Barsanti et al., 2009; Kurten et al., 2008). Recent experimental studies of nucleation have observed an enhancement in the ternary nucleation rate in the presence of amines (Berndt et al., 2010; Erupe et al., 2010). Amines have also been shown to contribute to particle growth (Wang et al., 2010a, b). We have determined the kinetics and thermodynamics of amine substitution for ammonia in positively-charged ammonium salt clusters of a wide array of sizes (Bzdek et al., 2010a, b, 2011). Regardless the type of salt, amine substitution is collision-limited on the cluster surface, which implies that small (1–2 nm diameter) ambient salts are likely to be aminium salts rather than ammonium salts, even if they were initially nucleated as ammonium salts.

Most recently, advances in ambient aerosol measurement instrumentation have permitted composition measurements of neutral and ionic molecular clusters. In the cluster chemical ionization mass spectrometer (cluster-CIMS), ambient neutral clusters are

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ionized by charge transfer from a reagent ion (Zhao et al., 2010). Clusters containing sulfuric acid are preferentially ionized by imparting a negative charge to them. The atmospheric pressure interface time-of-flight mass spectrometer (APi-TOF) samples ambient charged ions through an atmospheric pressure interface into a high resolution TOF mass analyzer (Junninen et al., 2010). Both of these methods ultimately perform analysis on a charged molecular cluster, whether that charge is intentionally imparted to the cluster (cluster-CIMS) or is naturally occurring (APi-TOF). However, the effect of charge on the composition of small molecular clusters is poorly understood.

This work examines the types of cluster distributions produced from electrospray ionization of ammonium sulfate and dimethylammonium sulfate solutions in both positive and negative polarity. The differences in cluster composition due to the polarity of charge, as well as their implications in terms of ambient measurements of molecular clusters, are discussed. Additionally, we determine both the kinetics of dimethylamine substitution for ammonia in negatively-charged ammonium bisulfate clusters and the kinetics of dimethylamine addition to these clusters in order to predict the composition of salt clusters likely to be present among ambient negative ions. The results of this work will facilitate interpretation of the mass spectra of ambient molecular clusters, as well as improve our understanding of the chemistry and composition of ca. 1–2 nm diameter particles.

2 Experimental section

The experimental setup was similar to that used in our previous work on positively-charged bisulfate, nitrate, and methanesulfonate clusters (Bzdek et al., 2010a, b, 2011). Charged ammonium bisulfate clusters were introduced into a 7T Bruker apex-Qe Fourier transform ion cyclotron resonance mass spectrometer (FTICR-MS) by electrospray of an ammonium sulfate (Aldrich) solution in 50/50 water/methanol. Charged dimethylammonium bisulfate clusters were produced by electrospray of a dimethylammonium sulfate solution made by combining solutions of dimethylamine (Fluka) in 50/50

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methanol/water and H₂SO₄ (Fisher) in 50/50 methanol/water. For broadband spectra, the solution concentration was 0.05 M, whereas for kinetics experiments a 1.0 mM solution was used.

Electrospray produced an array of clusters. For kinetics experiments, a specific cluster was isolated by mass selection and accumulation in a quadrupole. Ions were then transferred to the ICR cell where they were exposed to a constant pressure of reactant gas (dimethylamine: $2.4 \pm 0.5 \times 10^{-8}$ torr; ammonia: $1.0 \pm 0.2 \times 10^{-7}$ torr; Matheson Tri-Gas) that was introduced to the cell via a leak valve. Negatively-charged ammonium bisulfate clusters were exposed to dimethylamine gas in order to determine the kinetics of both substitution for ammonia and addition of dimethylamine to the clusters. Ammonium bisulfate clusters were also exposed to ammonia gas to determine the kinetics of ammonia addition to the clusters. Electrospray of a dimethylammonium sulfate solution in negative mode produced complicated spectra due to substitution of sodium for dimethylamine. Negatively-charged dimethylammonium bisulfate clusters were unstable upon mass selection; therefore, no kinetic data were obtained for exposure of these clusters to dimethylamine or ammonia gas.

A mass spectrum of ions in the ICR cell can be obtained at a specific trapping time. FTICR-MS provides high resolving power and high accuracy mass-to-charge (m/z) measurements, allowing for the assignment of unique elemental formulae to reactant and product ions. A plot of ion abundance as a function of trapping time (reaction profile) reveals the progress of the sequential reactions. In some of the more exothermic reactions, the reactant molecule induces a more complicated decomposition rather than substitution alone. A sensitivity analysis was performed in order to elucidate the effect of the decomposition pathways on measured kinetics (Bzdek et al., 2011). In the model, cluster decomposition was assumed to decrease the total cluster ion intensity by a factor similar to that observed experimentally, typically 20–30 % over a time period of about 1 s. The model showed that fitting the experimental data without taking cluster decomposition into account gave measured rate constants that were larger than the actual rate constants. Typically, the actual rate constants were about 30–40 % smaller

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than the measured values for the first or second reaction steps, but there was no difference between measured and actual rate constants for later reaction steps. Based on the measured rate constants in this work, it is unlikely that cluster decomposition significantly affected our kinetic measurements. All data were fit to the kinetic models using the simplex method of non-linear fitting embodied in the Solver function of the Microsoft Excel program. Second order rate constants were determined from gas pressure measurements in the ICR cell, as discussed elsewhere (Bzdek et al., 2010a, b); reactive uptake coefficients (reaction probabilities) were determined using capture collision theory to calculate the theoretical collisional rate constant (Ridge, 2003).

3 Results and discussion

3.1 Electro spray of ammonium sulfate and dimethylammonium sulfate solutions

Although the ions produced during electro spray ionization are not necessarily at equilibrium, the species most likely to receive a charge are those that tend to be more stable upon charging. In complex solutions containing several analytes, some chemical species will receive charge preferentially relative to other species in the same solution (Oss et al., 2010). Similar to this phenomenon, a solution of ammonium sulfate contains ammonium and sulfate ions that, upon ionization and transfer from the liquid phase to the gas phase, produce an array of unique cluster ions that contain differing numbers of ammonium and bisulfate ions. An analysis of the distribution of ionic species produced by electro spray of ammonium sulfate can give an indication of the relative stabilities of ammonium sulfate/ammonium bisulfate clusters that may be expected to be observed in ambient air.

Figure 1 provides mass spectra of an ammonium sulfate solution introduced by electro spray into the FTICR-MS operating in (a) positive and (b) negative modes. Interestingly, electro spray of the same solution produces cluster ions of very different

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ammonium sulfate eventually becomes preferred. For example, the elemental composition of 25 nm mass normalized diameter ammonium sulfate particles produced by electrospray gives a molar N:S ratio of 2, indicating these particles are composed of ammonium sulfate (N:S = 2), rather than ammonium bisulfate (N:S = 1) (Zordan et al., 2010).

Figure 2 presents a comparison of cluster acidity (defined as the number of unneutralized sulfuric acid molecules, n) in positively-charged and negatively-charged ammonium bisulfate and dimethylammonium bisulfate clusters. For a given type of cluster (ammonium bisulfate or dimethylammonium bisulfate) at a given polarity (positive or negative), the relative signal intensities of the different series of observed clusters produced by electrospray are plotted. Note that for positive ions, only singly-charged clusters are considered; for negative ions, only clusters containing ammonium or dimethylammonium are considered (sulfuric acid-bisulfate clusters are not included). For positive ions (Fig. 2a and b), clusters tend to be fully neutralized to bisulfate, as for both ammonium bisulfate and dimethylammonium bisulfate clusters the predominant cluster series observed in the mass spectrum is one that contains no unneutralized sulfuric acid ($n = 0$). However, the presence of dimethylammonium appears to enhance neutralization to bisulfate better than the presence of ammonium, as in Fig. 2b there is less signal intensity from the $[(\text{CH}_3)_2\text{NH}_2]_x(\text{HSO}_4)_{x-1}(\text{H}_2\text{SO}_4)]^+$ series than from the $[(\text{NH}_4)_x(\text{HSO}_4)_{x+1}(\text{H}_2\text{SO}_4)]^+$ series in Fig. 2a. The reason for this enhanced neutralization to bisulfate is likely because dimethylamine is a stronger base than ammonia.

Negative ions (Fig. 2c and d) show two interesting phenomena. The first observation is that negative ions in general differ significantly in composition relative to positive ions. The second observation is that the composition of ammonium bisulfate negative ions is different than the composition of dimethylammonium bisulfate negative ions, despite the similarity between the two in positive mode. Specifically, negative ions tend to be much more acidic (more unneutralized sulfuric acid) than positive ions. Whereas in positive mode, the predominant cluster series for both types of clusters contains no unneutralized sulfuric acid ($n = 0$), the most intense series of negatively-charged

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ammonium bisulfate clusters contains three unneutralized sulfuric acid molecules ($n = 3$; Fig. 2c) and the most intense series of negatively-charged dimethylammonium bisulfate clusters contains two unneutralized sulfuric acid molecules ($n = 2$; Fig. 2d). These observations also indicate that for negatively-charged clusters, dimethylammonium enhances neutralization of sulfuric acid better than ammonium. Again, this can be attributed to dimethylamine's stronger basicity.

Analysis of the ions produced from electrospray ionization of ammonium sulfate and dimethylammonium sulfate solutions in both positive and negative modes gives an indication of the most likely (most favorable) charged species to be formed. These observations may have important implications regarding the composition of ions in ambient air and will be discussed later.

3.2 Reactivity of negatively-charged ammonium bisulfate ions

In addition to examination of the composition of ions produced by electrospray of an ammonium sulfate solution, we examined the reactivity of these ions. We have previously studied the kinetics and thermodynamics of exposure of positively-charged ammonium bisulfate, ammonium nitrate, and ammonium methanesulfonate clusters to dimethylamine (Bzdek et al., 2010a, b, 2011). These previous studies focused solely on positive ions. Here, we discuss the reactivity of negatively-charged ammonium bisulfate clusters and sulfuric acid-bisulfate clusters upon exposure to dimethylamine gas.

The two series of clusters were analyzed in the kinetics experiments. The first series consisted of the sulfuric acid-bisulfate clusters ($[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_x]^-$), which occurred at low m/z values and contained no ammonium. The second series was the most abundant series of negative ions observed and was of the form $[(\text{NH}_4)_x(\text{HSO}_4)_{x+1}(\text{H}_2\text{SO}_4)_3]^-$. For these ammonium bisulfate clusters, three specific clusters were isolated and exposed to dimethylamine in order to elucidate the kinetics of amine substitution and amine addition to the cluster: a cluster containing one ammonium ion, $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_3]^-$; a cluster containing two ammonium

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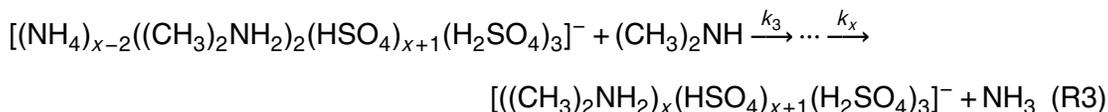
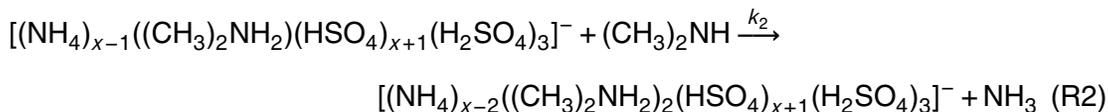
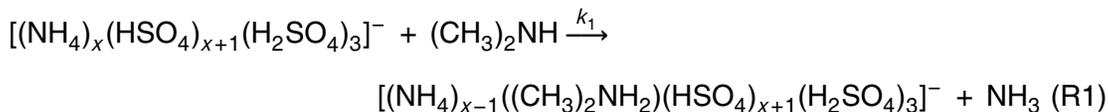
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ions, $[(\text{NH}_4)_2(\text{HSO}_4)_3(\text{H}_2\text{SO}_4)_3]^-$; and a cluster containing three ammonium ions, $[(\text{NH}_4)_3(\text{HSO}_4)_4(\text{H}_2\text{SO}_4)_3]^-$. Two unique chemical reactions occurred upon exposure of these clusters to dimethylamine: (1) dimethylamine substitution for ammonia and (2) dimethylamine addition to the cluster.

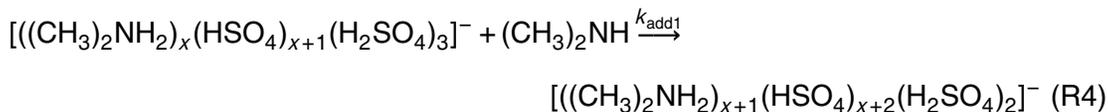
The substitution reaction can be described by Scheme 1. Note that in Scheme 1, the unneutralized sulfuric acid serves essentially as a spectator to the reaction.

SCHEME 1



Scheme 2 describes the addition reaction, whereby a dimethylamine can add to the cluster, effectively neutralizing the sulfuric acid and growing the cluster to a larger size.

SCHEME 2



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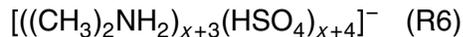
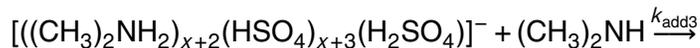
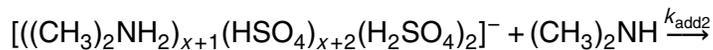
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5 The kinetics experiment is similar to our previous work on positive ions (Bzdek et al., 2010a, b, 2011). An ammonium bisulfate cluster is isolated by mass selection in the quadrupole of the mass spectrometer. The isolated cluster is then transferred to the ICR cell, where it is exposed to dimethylamine gas at constant pressure. Reaction time is varied, and time-dependent mass spectra are recorded. A reaction profile is created by plotting the relative abundance of each reaction step against exposure time. Pseudo-first order kinetics applies to these reactions because the pressure of the reactant gas is constant during the experiment. A fit of the reaction profile to pseudo-first order kinetics allows for quantitative determination of the rate constants for the sequential substitution and addition reactions.

15 Figure 3 presents the reaction profile (symbols) and statistical fit to pseudo-first order kinetics (lines) for $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_3]^-$ exposed to dimethylamine gas. At time $t = 0$ s, the initial cluster is the predominant ion observed, which is expected as no reaction should have occurred. With increasing exposure time to dimethylamine, ammonia is displaced by dimethylamine. Next, dimethylamine adds sequentially to the cluster, eventually neutralizing all sulfuric acid molecules to bisulfate. For the $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_3]^-$ cluster, the reactive uptake coefficient (reaction probability, γ) for substitution (i.e. Scheme 1) is $\gamma_{\text{sub}} = 0.24 \pm 0.07$. For the sequential addition steps (i.e. Scheme 2), the reactive uptake coefficients are $\gamma_{\text{add}1} = 0.13 \pm 0.04$, $\gamma_{\text{add}2} = 0.10 \pm 0.03$, and $\gamma_{\text{add}3} = 2.1 \pm 0.6 \times 10^{-2}$.

25 Table 1 presents pseudo-first order rate constants, second order rate constants, and reactive uptake coefficients for exposure of negatively-charged ammonium bisulfate

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clusters to dimethylamine. For clusters containing ammonium, substitution steps are fairly fast, with $\gamma_{\text{sub}} \geq 0.24$. These values for substitution are somewhat slower than those observed for positive ions, where γ_{sub} for ammonium ions on the cluster surface is typically greater than 0.50 (Bzdek et al., 2010a, b, 2011). The lower uptake coefficients are probably due to a decrease in the frequency of potentially reactive collisions (i.e. collision between the cluster and the amine would need to occur near the ammonium ion, but since fewer ammonium ions are in the cluster, there are more collisions where reaction would not occur). This helps to explain the observed increase in uptake coefficient for substitution as cluster size increases (and the number of reactive sites increases). The initial addition step is typically slower than the substitution steps by about a factor of two. The second addition step is slightly slower than the first. However, the final addition step (to neutralize all sulfuric acid to bisulfate) is significantly slower than the first two addition steps for all studied clusters. Future computational work investigating the thermodynamics and structure of these clusters would be beneficial in explaining these experimental observations. Based on the kinetic results, amine substitution for ammonia is facile. Addition is relatively fast and results in further neutralization of the cluster than what can be accomplished by just ammonium.

In addition to examining negatively-charged ammonium bisulfate clusters, we examined the kinetics of addition of dimethylamine to two sulfuric acid-bisulfate clusters, specifically $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$ and $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_2]^-$. Dimethylamine was not observed to add onto the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$ cluster ($\gamma_{\text{add}} < 9.7 \times 10^{-6}$); however dimethylamine did add to the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_2]^-$ cluster, though at a rate that was an order of magnitude slower than the slowest addition step in the ammonium-containing clusters ($\gamma_{\text{add}} = 1.0 \pm 0.3 \times 10^{-3}$). These observations suggest that for negatively-charged sulfuric acid-bisulfate clusters in this size range, the incorporation of amine is not kinetically favorable. These observations agree with recent computational modeling that suggests similar negatively-charged sulfuric acid-bisulfate clusters containing dimethylamine are not thermodynamically stable (Kurten et al., 2011).

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The reactivity of these same clusters was examined upon exposure to ammonia gas. Table 3 provides pseudo-first order rate constants, second order rate constants, and reactive uptake coefficients for exposure of the three ammonium bisulfate clusters and the two sulfuric acid-bisulfate clusters to ammonia. Ammonia added to neither the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$ cluster ($\gamma_{\text{add}} < 8.1 \times 10^{-5}$) nor the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_2]^-$ cluster ($\gamma_{\text{add}} < 5.3 \times 10^{-4}$). For ammonium-containing clusters, ammonia did not add onto the $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_3]^-$ cluster ($\gamma_{\text{add}} < 9.9 \times 10^{-4}$), but did add to the $[(\text{NH}_4)_2(\text{HSO}_4)_3(\text{H}_2\text{SO}_4)_3]^-$ and $[(\text{NH}_4)_3(\text{HSO}_4)_4(\text{H}_2\text{SO}_4)_3]^-$ clusters ($\gamma_{\text{add}} = 5.1 \pm 1.5 \times 10^{-3}$ and $1.2 \pm 0.4 \times 10^{-2}$, respectively). Addition was slow, so only the first addition step was observed. The addition rates for ammonia to these clusters are more than one order of magnitude slower than those for addition of dimethylamine to the dimethylammonium-substituted clusters. The measured kinetics for ammonia addition to the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$ and $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_2]^-$ clusters compares favorably with quantum chemical calculations performed previously on these clusters. Ortega et al. (2008) performed calculations that suggest the presence of ammonia is unfavorable in the $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$ cluster, which agrees with our kinetics observations, as the uptake coefficient was $< 8.1 \times 10^{-5}$. They also predict that the favorability for ammonia in these clusters increases with increasing cluster size, as we have demonstrated here.

4 Atmospheric implications

This study examined the distribution of ionic species produced by electrospray of ammonium sulfate in both positive and negative polarities. The ions produced by electrospray give an indication of what ionic species might be expected in ambient air. Recent advances in instrumentation to measure charged clusters in ambient air permit comparison of the electrospray ionization mass spectra to ambient ion spectra. Specifically, the API-TOF permits detection of either positively- or negatively-charged clusters up to 2000 m/z (Junninen et al., 2010; Ehn et al., 2010). API-TOF analysis

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of nebulized sulfuric acid mixed with laboratory air yielded a spectrum with some prominent peaks at m/z values that correspond to those dominant negatively-charged ammonium bisulfate series observed from electrospray of an ammonium sulfate solution (Junninen et al., 2010). Ambient measurements of negative ions $< 500 m/z$ during the daytime are dominated by sulfuric acid-bisulfate clusters (e.g. $[\text{HSO}_4]^-$, $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)]^-$, $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_2]^-$, and $[(\text{HSO}_4)(\text{H}_2\text{SO}_4)_3]^-$). Nonetheless, signal from negatively-charged clusters containing ammonium has also been detected (e.g. $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_2]^-$) (Ehn et al., 2010). These ambient observations confirm the hypothesis that bases will not be observed in low m/z negative ion clusters.

For negative ions, amine-ammonia chemistry will only be evident and important in larger cluster ions ($> 400 m/z$), which appear to be much lower in concentration than the low m/z cluster ions and therefore harder to detect. However, as illustrated in Fig. 1, amine-ammonia chemistry is likely to influence even the smallest positive ions. Therefore, these types of clusters should be more evident in ambient positive ion spectra than in ambient negative ion spectra. Ambient positive ion spectra, though, are quite complex (Ehn et al., 2010; Junninen et al., 2010), as there are many other chemical species in the atmosphere competing for a limited amount of charge. Observation of positively-charged ammonium-ammonium bisulfate cluster ions may be obscured by competition with other (non-salt) species for charge or even simply by the vast number of positively-charged ambient species.

The results of our kinetics experiments have important atmospheric implications, especially since many of these charged clusters may be present in ambient air. Charge plays a critical role in the formation of the initial reactive clusters produced by the electrospray process and may affect the reaction rates for the smallest cluster sizes. However, the role of charge diminishes with increasing cluster size. As discussed previously, for both positive and negative ions, the composition appears to converge to ammonium bisulfate at large cluster sizes. Whereas amine substitution for ammonia may or may not be important to the nucleation of new particles, it is likely to impact particle growth in the 1–2 nm size range. Because substitution is nearly collision-limited for

the negatively-charged ammonium bisulfate clusters, and because these results agree very well with those of positive ions (Bzdek et al., 2010a, b, 2011), ammonium salt clusters in the 1–2 nm size range whether charged or not would be expected to quickly react with amine to form aminium salts upon atmospheric collision of the ammonium salt with the gas-phase amine. Finally, the kinetics of ammonia addition to these clusters is 1–2 orders of magnitude slower than both mechanisms of dimethylamine incorporation into the cluster (substitution and addition). Therefore, when ambient concentrations of ammonia and amine are comparable, amine chemistry will likely compete favorably with ammonia chemistry.

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Table 1. Pseudo-first order rate constants, second order rate constants, and uptake coefficients for reaction of dimethylamine with negatively-charged ammonium bisulfate clusters.^a

Cluster	Reaction step	Pseudo-first order rate constant (s ⁻¹)	Second order rate constant (cm ³ molecule ⁻¹ s ⁻¹)	Uptake coefficient, γ
[(HSO ₄)(H ₂ SO ₄)] ⁻	Addition	< 1.1 × 10 ⁻⁵	< 1.3 × 10 ⁻¹⁴	< 9.7 × 10 ⁻⁶
[(HSO ₄) ₂ (H ₂ SO ₄) ₂] ⁻	Addition	1.2 ± 0.2 × 10 ⁻³	1.4 ± 0.3 × 10 ⁻¹²	1.0 ± 0.3 × 10 ⁻³
[(NH ₄)(HSO ₄) ₂ (H ₂ SO ₄) ₃] ⁻	Substitution	0.26 ± 0.03	3.1 ± 0.6 × 10 ⁻¹⁰	0.24 ± 0.07
	Addition 1	0.14 ± 0.01	1.7 ± 0.3 × 10 ⁻¹⁰	0.13 ± 0.04
	Addition 2	0.11 ± 0.01	1.3 ± 0.3 × 10 ⁻¹⁰	0.10 ± 0.03
	Addition 3	2.2 ± 0.3 × 10 ⁻²	2.6 ± 0.5 × 10 ⁻¹¹	2.1 ± 0.6 × 10 ⁻²
[(NH ₄) ₂ (HSO ₄) ₃ (H ₂ SO ₄) ₃] ⁻	Substitution 1	0.48 ± 0.08	5.6 ± 1.1 × 10 ⁻¹⁰	0.44 ± 0.13
	Substitution 2	0.38 ± 0.06	4.5 ± 0.9 × 10 ⁻¹⁰	0.36 ± 0.11
	Addition 1	0.19 ± 0.03	2.2 ± 0.4 × 10 ⁻¹⁰	0.17 ± 0.05
	Addition 2	0.13 ± 0.02	1.6 ± 0.3 × 10 ⁻¹⁰	0.13 ± 0.04
	Addition 3	3.6 ± 0.5 × 10 ⁻²	4.2 ± 0.8 × 10 ⁻¹¹	3.4 ± 1.0 × 10 ⁻²
[(NH ₄) ₃ (HSO ₄) ₄ (H ₂ SO ₄) ₃] ⁻	Substitution 1	0.56 ± 0.06	6.6 ± 1.3 × 10 ⁻¹⁰	0.53 ± 0.16
	Substitution 2	0.51 ± 0.07	6.0 ± 1.2 × 10 ⁻¹⁰	0.48 ± 0.14
	Substitution 3	0.43 ± 0.06	5.1 ± 1.0 × 10 ⁻¹⁰	0.41 ± 0.12
	Addition 1	0.23 ± 0.02	2.3 ± 0.5 × 10 ⁻¹⁰	0.22 ± 0.07
	Addition 2	0.13 ± 0.01	1.2 ± 0.2 × 10 ⁻¹⁰	0.12 ± 0.04
	Addition 3	7.0 ± 0.7 × 10 ⁻²	6.8 ± 1.4 × 10 ⁻¹¹	6.7 ± 2.0 × 10 ⁻²

^a $P_{\text{dimethylamine}}$: 2.4 ± 0.5 × 10⁻⁸ torr

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Table 2. Pseudo-first order rate constants, second order rate constants, and uptake coefficients for reaction of ammonia with negatively-charged ammonium bisulfate clusters.^a

Cluster	Reaction step	Pseudo-first order rate constant (s ⁻¹)	Second order rate constant (cm ³ molecule ⁻¹ s ⁻¹)	Uptake coefficient, γ
[(HSO ₄)(H ₂ SO ₄)] ⁻	Addition	< 5.1 × 10 ⁻⁴	< 1.6 × 10 ⁻¹³	< 8.1 × 10 ⁻⁵
[(HSO ₄)(H ₂ SO ₄) ₂] ⁻	Addition	< 3.3 × 10 ⁻³	< 1.1 × 10 ⁻¹²	< 5.3 × 10 ⁻⁴
[(NH ₄)(HSO ₄) ₂ (H ₂ SO ₄) ₃] ⁻	Addition	< 6.1 × 10 ⁻³	< 1.9 × 10 ⁻¹²	< 9.9 × 10 ⁻⁴
[(NH ₄) ₂ (HSO ₄) ₃ (H ₂ SO ₄) ₃] ⁻	Addition	3.1 ± 0.1 × 10 ⁻²	1.0 ± 0.2 × 10 ⁻¹¹	5.1 ± 1.5 × 10 ⁻³
[(NH ₄) ₃ (HSO ₄) ₄ (H ₂ SO ₄) ₃] ⁻	Addition	7.1 ± 0.2 × 10 ⁻²	2.3 ± 0.5 × 10 ⁻¹¹	1.2 ± 0.4 × 10 ⁻²

^a P_{ammonia} : 1.0 ± 0.2 × 10⁻⁷ torr

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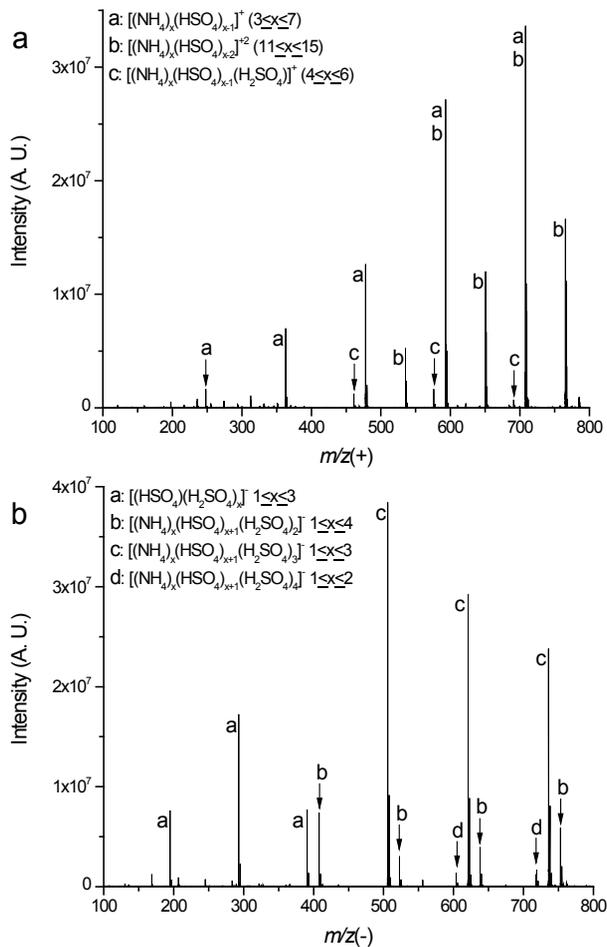


Fig. 1. Electrospray ionization mass spectra of an ammonium sulfate solution in **(a)** positive and **(b)** negative modes.

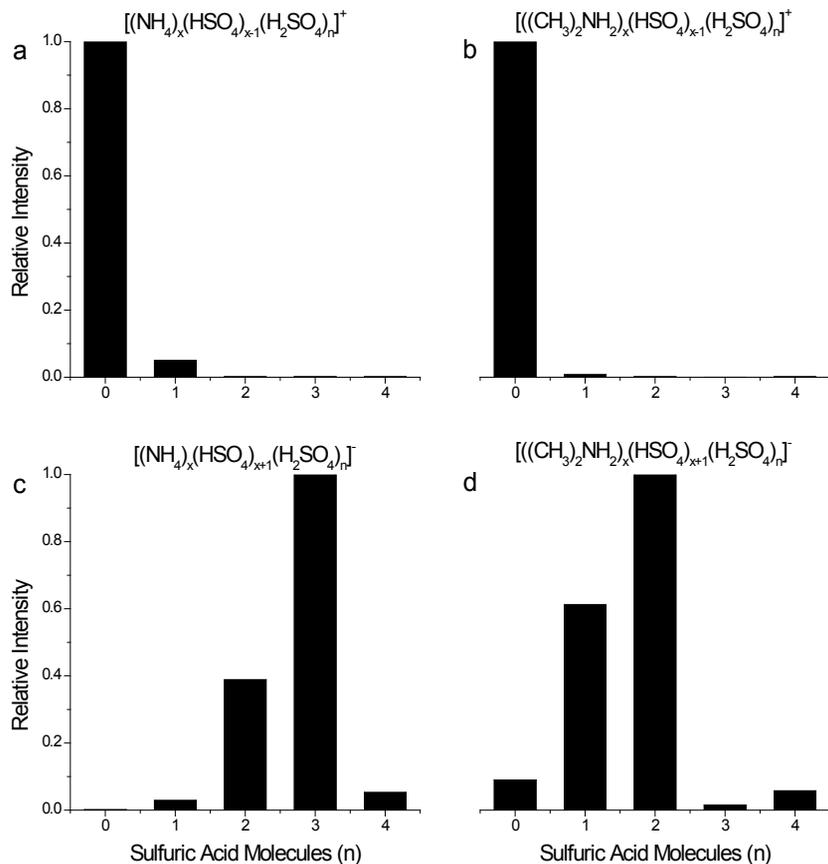


Fig. 2. Cluster acidity (defined as the number of unneutralized sulfuric acid molecules, n , in a cluster) for clusters formed through electrospray of ammonium sulfate and dimethylammonium sulfate solutions in positive and negative polarities.

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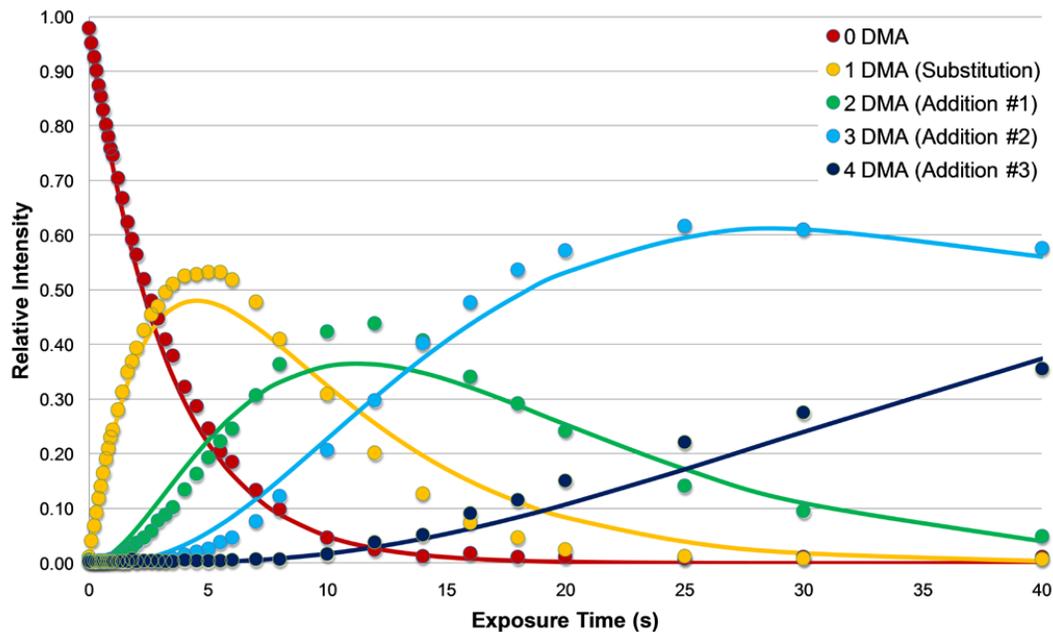


Fig. 3. Reaction profile (symbols) and statistical fit to pseudo-first order kinetics (lines) for exposure of $[(\text{NH}_4)(\text{HSO}_4)_2(\text{H}_2\text{SO}_4)_3]^-$ to dimethylamine gas ($P_{\text{dimethylamine}}: 2.4 \pm 0.5 \times 10^{-8}$ torr).

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