ELECTRONIC SUPPLEMENT FOR
Modeling chemistry in and above snow at Summit, Greenland
Part 1: Model description and results

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Jennie L. Thomas[1], Jochen Stutz[1], Barry Lefer[2], L. Gregory Huey[3], Kenjiro Toyota[4,5], Jack E. Dibb[6], Roland von Glasow[7]

[1]University of California, Los Angeles; Department of Atmospheric and Oceanic Sciences, Los Angeles, CA, USA,
[2]Earth and Atmospheric Sciences Department, University of Houston, Houston, TX, USA,
[3]School of Earth and Atmospheric Sciences, Georgia Institute of Technology, Atlanta, Georgia, USA,
[4]Department of Earth and Space Science and Engineering, York University, Toronto, Ontario, Canada,
[5]Air Quality Research Division, Science and Technology Branch, Environment Canada, Toronto, Ontario, Canada,
[6]Institute for the Study of Earth, Oceans and Space, University of New Hampshire, Durham, New Hampshire, USA,
[7]School of Environmental Sciences, University of East Anglia, Norwich, UK

SUPPLEMENT INCLUDES:
Gas and aqueous species included in the model, henry constants, accommodation coefficients, gas phase reaction rates, and aqueous phase reaction rates.

Table 1: A complete list of all gas and aqueous phase species included in the MISTRA-SNOW model.

Gas phase
$O(^1D)$, O_2, O_3, OH, HO_2, H_2O_2, H_2O
NO, NO_2, NO_3, N_2O_5, HONO, HNO_3, HNO_4, PAN, NH_3
CO, CO_2, CH_4, C_2H_6, C_2H_4, HCHO, HCOOH, ALD (i.e., CH_3CHO), CH_2O_2, $HOCH_2O_2$, CH_3CO_3, CH_3O_2, $C_2H_5O_2$, CH_3O_2, EO_2 (i.e., $H_2C(OH)CH_2OO$), CH_2O_2, ROOH (i.e., alkylhydroperoxides)
SO_2, SO_3, $HOSO_2$, H_2SO_4, DMS, CH_3SCH_2OO, DMSO, $DMSO_2$, CH_3S, CH_3SO, CH_3SO_2, CH_3SO_3, CH_3SO_2H, CH_3SO_3H
Cl, ClO, OClO, HCl, HOCl, Cl_2, Cl_2O_2 $ClNO_2$, $ClNO_3$
Br, BrO, HBr, HOBr, Br_2, $BrNO_2$, $BrNO_3$, BrCl
Liquid phase (neutrals)
$O(^3P)$, O_2, O_3, OH, HO_2, H_2O_2, H_2O
NO, NO_2, NO_3, HONO, HNO_3, HNO_4, NH_3
CO_2, HCHO, HCOOH, CH_3OH, CH_3OO, CH_3OOH, DOM
SO_2, H_2SO_4, DMSO, $DMSO_2$, CH_3SO_2H, CH_3SO_3H
Cl, HCl, HOCl, Cl_2
Br, HBr, HOBr, Br_2, BrCl
Liquid phase (ions)
H^+, OH^-, O_2^-
NO_2^-, NO_3^-, NO_4^-, NH_4^+
HCO_3^-, CO_3^-, $HCOO^-$
HSO_3^-, $SO_3{}^2-$, HSO_4^-, SO_4^{2-}, HSO_5^-, SO_3^-, SO_4^-, SO_5^-, $CH_3SO_3^-$, $CH_2OHSO_2^-$, $CH_2OHSO_3^-$
Cl^-, Cl_2^-, ClO^-, $ClOH^-$
Br^-, Br_2^-, BrO^-, $BrCl_2^-$, Br_2Cl^-, $BrOH^-$

Table 2: Henry constants and accommodation coefficients.[‡]

species	K_H^0 [M/atm]	$-\Delta_{soln}H/R$ [K]	reference	α^0	$-\Delta_{obs}H/R$ [K]	reference
O_3	1.2×10^{-2}	2560	(Chameides, 1984)	0.04	(water ice at 195-262 K)	(Sander et al., 2006)
O_2	1.3×10^{-3}	1500	(Wilhelm et al., 1977)	0.01	2000	estimated
OH	3.0×10^{1}	4300	(Hanson et al., 1992)	0.1	(water ice at 205-253 K)	(Sander et al., 2006)
HO_2	3.9×10^{3}	5900	(Hanson et al., 1992)	0.02	(at 275 K)	(Sander et al., 2006)
H_2O_2	1.0×10^{5}	6338	(Lind and Kok, 1994)	0.077	2769	(Worsnop et al., 1989)
NO	1.9×10^{-3}	1480	(Schwartz and White, 1981)	5×10^{-5}	0	(Saastad et al., 1993)
NO_2	6.4×10^{-3}	2500	(Lelieveld and Crutzen, 1991)	0.0001	(water ice at 195 K)	(Sander et al., 2006)
NO_3	2.0	2000	(Thomas et al., 1993)	0.04	(at 273 K)	(Rudich et al., 1996)
N_2O_5	∞	—		0.1	(at 195-300 K)	(DeMore et al., 1997)
HONO	4.9×10^{1}	4780	(Schwartz and White, 1981)	0.001	(water ice at 180-200 K)	(Sander et al., 2006)
HNO_3	1.7×10^{5}	8694	(Lelieveld and Crutzen, 1991)	0.003	(water ice at 220 K)	(Sander et al., 2006)
HNO_4	1.2×10^{4}	6900	(Régimbal and Mozurkewich, 1997)	0.1	(at 200 K)	(DeMore et al., 1997)
NH_3	5.8×10^{1}	4085	(Chameides, 1984)	0.06	(at 295 K)	(DeMore et al., 1997)
CH3OO	6.0	$=HO_2$	(Pandis and Seinfeld, 1989)	0.01	2000	estimated
ROOH	3.0×10^{2}	5322	(Lind and Kok, 1994)	0.0046	3273	(Magi et al., 1997)
HCHO	7.0×10^{3}	6425	(Chameides, 1984)	0.04	(at 260-270 K)	(DeMore et al., 1997)
HCOOH	3.7×10^{3}	5700	(Chameides, 1984)	0.014	3978	(DeMore et al., 1997)
CO_2	3.1×10^{-2}	2423	(Chameides, 1984)	0.01	2000	estimated
HCl	1.2	9001	(Brimblecombe and Clegg, 1989)	0.3	water ice (191 211 K) α = 0.18 for liquid water at 273 K	(Sander et al., 2006)
HOCl	6.7×10^{2}	5862	(Huthwelker et al., 1995)	=HOBr	=HOBr	estimated
$ClNO_3$	∞	—		0.1	(at RT)	(Koch and Rossi, 1998)
Cl_2	9.1×10^{-2}	2500	(Wilhelm et al., 1977)	0.0001	(water ice at 200 K)	(Sander et al., 2006)
HBr	1.3	10239	(Brimblecombe and Clegg, 1989)	0.2	(water ice at 200 K)	(Sander et al., 2006)
HOBr	9.3×10^{1}	=HOCl	(Vogt et al., 1996)	0.003	(water ice at 223-239 K)	(Sander et al., 2006)
$BrNO_3$	∞	—		0.8	0	(Hanson et al., 1996)
Br_2	7.6×10^{-1}	4094	(Dean, 1992)	0.038	6546	(Hu et al., 1995)
BrCl	9.4×10^{-1}	5600	(Bartlett and Margerum, 1999)	0.15	(at 270-285 K)	(Sander et al., 2006)

[‡] For ROOH the values of CH_3OOH have been assumed. The temperature dependence is for the Henry constants is $K_H = K_H^0 \times \exp(\frac{-\Delta_{soln}H}{R}(\frac{1}{T} - \frac{1}{T_0}))$, $T_0 = 298$ K and for the accommodation coefficients $dln(\frac{\alpha}{1-\alpha})/d(\frac{1}{T}) = \frac{-\Delta_{obs}H}{R}$. RT stands for "room temperature".

Table 2 - Henry constants and accommodation coefficients.

species	K_H^0 [M/atm]	$-\Delta_{soln}H/R$ [K]	reference	α^0	$-\Delta_{obs}H/R$ [K]	reference
DMS	4.8×10^{-1}	3100	(De Bruyn et al., 1995)	0.01		assumed
DMSO	5.0×10^4	=HCHO	(De Bruyn et al., 1994)	0.048	2578	(De Bruyn et al., 1994)
$DMSO_2$	∞	—	assumed	0.03	5388	(De Bruyn et al., 1994)
SO_2	1.2	3120	(Chameides, 1984)	0.11	0	(DeMore et al., 1997)
H_2SO_4	∞	—		0.65	(at 303 K)	(Pöschl et al., 1998)
CH_3SO_2H	∞	—	assumed	0.0002	0	(Lucas and Prinn, 2002)
CH_3SO_3H	∞	—	assumed	0.076	1762	(De Bruyn et al., 1994)
CH_4	1.3×10^{-3}	—	(Mackay and Shiu., 1981)	0.1	0	assumed

‡For ROOH the values of CH_3OOH have been assumed. The temperature dependence is for the Henry constants is $K_H = K_H^0 \times \exp(\frac{-\Delta_{soln}H}{R}(\frac{1}{T} - \frac{1}{T_0}))$, $T_0 = 298$ K and for the accommodation coefficients $dln(\frac{\alpha}{1-\alpha})/d(\frac{1}{T}) = \frac{-\Delta_{obs}H}{R}$. RT stands for “room temperature”.

Table 2: Gas phase reactions.

no	reaction	n	A [(cm^{-3})$^{1-n}$s^{-1}]	$-E_a$ / R [K]	reference
Ox and HOx reactions					
O1	$O(^1D) + O_2 \longrightarrow O_3$	2	3.2×10^{-11}	70	Atkinson et al. (2004)
O2	$O(^1D) + N_2 \longrightarrow O_3$	2	1.8×10^{-11}	110	Atkinson et al. (2004)
O3	$O(^1D) + H_2O \longrightarrow 2\,OH$	2	2.2×10^{-10}		Atkinson et al. (2004)
O4	$OH + O_3 \longrightarrow HO_2 + O_2$	2	1.7×10^{-12}	-940	Atkinson et al. (2004)
O5	$OH + HO_2 \longrightarrow H_2O + O_2$	2	4.8×10^{-11}	250	Atkinson et al. (2004)
O6	$OH + H_2O_2 \longrightarrow HO_2 + H_2O$	2	2.9×10^{-12}	-160	Atkinson et al. (2004)
O7	$HO_2 + O_3 \longrightarrow OH + 2\,O_2$	2	1.0×10^{-14}	-490	Atkinson et al. (2004)
O8	$HO_2 + HO_2 \longrightarrow H_2O_2 + O_2$	2	[b]		Atkinson et al. (2006)
O9	$O_3 + h\nu \longrightarrow O_2 + O(^1D)$	1	[a]		DeMore et al. (1997)
O10	$H_2O_2 + h\nu \longrightarrow 2\,OH$	1	[a]		DeMore et al. (1997)
NOy reactions					
N1	$NO + OH \xrightarrow{M} HONO$	3	[b]		Sander et al. (2003)
N2	$NO + HO_2 \longrightarrow NO_2 + OH$	2	3.5×10^{-12}	250	Atkinson et al. (2004)
N3	$NO + O_3 \longrightarrow NO_2 + O_2$	2	3.0×10^{-12}	-1500	Sander et al. (2003)
N4	$NO + NO_3 \longrightarrow 2\,NO_2$	2	1.5×10^{-11}	170	Sander et al. (2003)
N5	$NO_2 + OH \xrightarrow{M} HNO_3$	3	[b]		Sander et al. (2003)
N6	$NO_2 + HO_2 \xrightarrow{M} HNO_4$	3	[b]		Atkinson et al. (2004)
N7	$NO_2 + O_3 \longrightarrow NO_3 + O_2$	2	1.2×10^{-13}	-2450	Sander et al. (2003)
N8	$NO_2 + h\nu \longrightarrow NO + O_3$	1	[a]		DeMore et al. (1997)
N9	$NO_2 + NO_3 \xrightarrow{M} N_2O_5$	3	[b]		Sander et al. (2003)
N10	$NO_3 + h\nu \longrightarrow NO + O_2$	1	[a]		Wayne et al. (1991)
N11	$NO_3 + HO_2 \longrightarrow 0.3\ HNO_3 + 0.7\ OH + 0.7\ NO_2 + O_2$	2	4.0×10^{-12}		Atkinson et al. (2004)
N12	$NO_3 + NO_3 \longrightarrow NO_2 + NO_2 + O_2$	2	8.5×10^{-13}	-2450	Sander et al. (2003)
N13	$NO_3 + h\nu \longrightarrow NO_2 + O_3$	1	[a]		Wayne et al. (1991)
N14	$N_2O_5 \xrightarrow{M} NO_2 + NO_3$	2	[b]		Sander et al. (2003)
N15	$N_2O_5 + H_2O \longrightarrow 2\,HNO_3$	2	2.6×10^{-22}		Atkinson et al. (2004)
N16	$N_2O_5 + h\nu \longrightarrow NO_2 + NO_3$	1	[a]		DeMore et al. (1997)
N17	$HONO + OH \longrightarrow NO_2$	2	1.8×10^{-11}	-390	Sander et al. (2003)
N18	$HONO + h\nu \longrightarrow NO + OH$	1	[a]		DeMore et al. (1997)
N19	$HNO_3 + h\nu \longrightarrow NO_2 + OH$	1	[a]		DeMore et al. (1997)
N20	$HNO_3 + OH \longrightarrow NO_3 + H_2O$	2	[b]		Atkinson et al. (2004)
N21	$HNO_4 \xrightarrow{M} NO_2 + HO_2$	2	[b]		Sander et al. (2003)
N22	$HNO_4 + OH \longrightarrow NO_2 + H_2O + O_2$	2	1.3×10^{-12}	380	Haggerstone et al. (2005)
N23	$HNO_4 + h\nu \longrightarrow NO_2 + HO_2$	1	[a]		DeMore et al. (1997)
N24	$HNO_4 + h\nu \longrightarrow OH + NO_3$	1	[a]		DeMore et al. (1997)
N25	$RONO_2 + OH \longrightarrow H_2O + NO_2$	2	1.3×10^{-12}		pers. comm. with R. Sander
N26	$RONO_2 + h\nu \longrightarrow NO_2$	1	[a]		assumed similar to HNO_3 photolysis
organic reactions					
C1	$CO + OH \xrightarrow{O_2} HO_2 + CO_2$	2	[b]		Sander et al. (2003)
C2	$CH_4 + OH \xrightarrow{O_2} CH_3O_2 + H_2O$	2	2.4×10^{-12}	-1775	Sander et al. (2003)
C3	$C_2H_6 + OH \longrightarrow C_2H_5O_2 + H_2O$	2	1.7×10^{-11}	-1232	Lurmann et al. (1986)
C4	$C_2H_4 + OH \longrightarrow C_2H_4OHO_2$	2	1.66×10^{-12}	474	Lurmann et al. (1986), see note
C5	$C_2H_4 + O_3 \longrightarrow HCHO + 0.4\ CH_2O_2 + 0.12\ HO_2 + 0.42\ CO + 0.06CH_4$	2	1.2×10^{-14}	-2633	Lurmann et al. (1986), see note
C6	$HO_2 + CH_3O_2 \longrightarrow ROOH + O_2$	2	4.1×10^{-13}	750	Sander et al. (2003)
C7	$HO_2 + C_2H_5O_2 \longrightarrow ROOH + O_2$	2	7.5×10^{-13}	700	Sander et al. (2003)
C8	$HO_2 + CH_3CO_3 \longrightarrow ROOH + O_2$	2	4.5×10^{-13}	1000	DeMore et al. (1997)
C9	$CH_3O_2 + CH_3O_2 \longrightarrow 1.4\ HCHO + 0.8\ HO_2 + O_2$	2	1.5×10^{-13}	220	Lurmann et al. (1986)

n reaction order, [a] photolysis rates calculated online, [b] special rate functions.

Table 2 - Gas phase reactions.

no	reaction	n	A [$(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}$]	$-E_a$ / R [K]	reference
C10	$C_2H_5O_2 + NO \longrightarrow ALD + HO_2 + NO_2$	2	4.2×10^{-12}	180	Lurmann et al. (1986)
C11	$2\,C_2H_5O_2 \longrightarrow 1.6\ ALD + 1.2\ HO_2$	2	5.0×10^{-14}		Lurmann et al. (1986)
C12	$C_2H_4OHO_2 + NO \longrightarrow NO_2 + 2\,HCHO + HO_2$	2	4.2×10^{-12}	180	Lurmann et al. (1986)
C13	$C_2H_4OHO_2 + C_2H_4OHO_2 \longrightarrow 2.4\ HCHO + 1.2\ HO_2 + 0.4\ ALD$	2	5.0×10^{-14}		Lurmann et al. (1986)
C14	$HO_2 + C_2H_4OHO_2 \longrightarrow ROOH + O_2$	2	3.0×10^{-12}		Lurmann et al. (1986)
C15	$HCHO + h\nu \longrightarrow 2\,HO_2 + CO$	1	^{a}		DeMore et al. (1997)
C16	$HCHO + h\nu \longrightarrow CO + H_2$	1	^{a}		DeMore et al. (1997)
C17	$HCHO + OH \xrightarrow{O_2} HO_2 + CO + H_2O$	2	1.0×10^{-11}		DeMore et al. (1997)
C18	$NO_3 + HCHO \xrightarrow{O_2} HNO_3 + HO_2 + CO$	2	5.8×10^{-16}		DeMore et al. (1997)
C19	$ALD + OH \longrightarrow CH_3CO_3 + H_2O$	2	6.9×10^{-12}	250	Lurmann et al. (1986)
C20	$ALD + NO_3 \longrightarrow HNO_3 + CH_3CO_3$	2	1.40×10^{-15}		DeMore et al. (1997)
C21	$ALD + h\nu \longrightarrow CH_3O_2 + HO_2 + CO$	1	^{a}		Lurmann et al. (1986)
C22	$ALD + h\nu \longrightarrow CH_4 + CO$	1	^{a}		Lurmann et al. (1986)
C23	$HOCH_2O_2 + NO \longrightarrow HCOOH + HO_2 + NO_2$	2	4.2×10^{-12}	180	Lurmann et al. (1986)
C24	$HOCH_2O_2 + HO_2 \longrightarrow HCOOH + H_2O + O_2$	2	2.00×10^{-12}		Lurmann et al. (1986)
C25	$2\,HOCH_2O_2 \longrightarrow 2\,HCOOH + 2\,HO_2 + 2\,O_2$	2	1.0×10^{-13}		Lurmann et al. (1986)
C26	$HCOOH + OH \xrightarrow{O_2} HO_2 + H_2O + CO_2$	2	4.0×10^{-13}		DeMore et al. (1997)
C27	$CH_3CO_3 + NO_2 \longrightarrow PAN$	2	4.7×10^{-12}		Lurmann et al. (1986)
C28	$PAN \longrightarrow CH_3CO_3 + NO_2$	1	1.9×10^{16}	-13543	DeMore et al. (1997)
C29	$CH_3CO_3 + NO \longrightarrow CH_3O_2 + NO_2 + CO_2$	2	4.2×10^{-12}	180	Lurmann et al. (1986)
C30	$CH_3O_2 + NO \xrightarrow{O_2} HCHO + NO_2 + HO_2$	2	3.0×10^{-12}	280	DeMore et al. (1997)
C31	$ROOH + OH \longrightarrow 0.7CH_3O_2 + 0.3HCHO + 0.3OH$	2	3.8×10^{-12}	200	DeMore et al. (1997), see note
C32	$ROOH + h\nu \longrightarrow HCHO + OH + HO_2$	1	^{a}		DeMore et al. (1997), see note
	S reactions				
S1	$SO_2 + OH \xrightarrow{M} HOSO_2$	3	^{b}		Atkinson et al. (2004)
S2	$HOSO_2 + O_2 \longrightarrow HO_2 + SO_3$	2	1.3×10^{-12}	-330	Atkinson et al. (2004)
S3	$SO_3 \xrightarrow{H_2O} H_2SO_4$	1	^{b}		Jayne et al. (1997)
S4	$DMS + OH \xrightarrow{O_2} DMOO + H_2O$	2	^{b}		Atkinson et al. (1997)
S5	$DMS + OH \xrightarrow{O_2} DMSO + HO_2$	2	^{b}		Atkinson et al. (1997)
S6	$DMS + NO_3 \xrightarrow{O_2} DMOO + HNO_3$	2	1.9×10^{-13}	520	Atkinson et al. (1999)
S7	$DMS + Cl \xrightarrow{O_2} DMOO + HCl$	2	3.3×10^{-10}		Jefferson et al. (1994)
S8	$DMS + Br \xrightarrow{O_2} DMOO + HBr$	2	9.0×10^{-11}	-2386	Ingham et al. (1999)
S9	$DMS + BrO \longrightarrow DMSO + Br$	2	2.54×10^{-14}	850	Ingham et al. (1999)
S10	$DMS + ClO \longrightarrow DMSO + Cl$	2	9.5×10^{-15}		Barnes et al. (1991)
S11	$DMOO + NO \longrightarrow HCHO + CH_3S + NO_2$	2	4.9×10^{-12}	263	Urbanski et al. (1997)
S12	$DMOO + DMOO \xrightarrow{O_2} 2\,HCHO + 2\,CH_3S$	2	1.0×10^{-11}		Urbanski et al. (1997); Atkinson et al. (2004)
S13	$CH_3S + O_3 \longrightarrow CH_3SO + O_2$	2	1.15×10^{-12}	432	Atkinson et al. (2004)
S14	$CH_3S + NO_2 \longrightarrow CH_3SO + NO$	2	3.0×10^{-11}	210	Atkinson et al. (2004)
S15	$CH_3SO + NO_2 \xrightarrow{O_2} 0.82CH_3SO_2 + 0.18SO_2 + 0.18CH_3O_2 + NO$	2	1.2×10^{-11}		Atkinson et al. (2004); Kukui et al. (2000), product ratios from van Dingenen et al. (1994)
S16	$CH_3SO + O_3 \longrightarrow CH_3SO_2 + O_2$	2	6.0×10^{-13}		Atkinson et al. (2004)
S17	$CH_3SO_2 \xrightarrow{O_2} CH_3O_2 + SO_2$	1	1.36×10^{14}	-8656	Kukui et al. (2000)
S18	$CH_3SO_2 + NO_2 \longrightarrow CH_3SO_3 + NO$	2	2.2×10^{-12}		Ray et al. (1996)
S19	$CH_3SO_2 + O_3 \longrightarrow CH_3SO_3$	2	5.0×10^{-15}		Ray et al. (1996)
S20	$CH_3SO_3 + HO_2 \longrightarrow CH_3SO_3H$	2	5.0×10^{-11}		Barone et al. (1995)
S21	$CH_3SO_3 \xrightarrow{H_2O,\ O_2} CH_3O_2 + H_2SO_4$	1	1.36×10^{14}	-11071	Barone et al. (1995)

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions.

Table 2 - Gas phase reactions.

no	reaction	n	A [(cm^{-3})$^{1-n}$s^{-1}]	$-E_a$ / R [K]	reference
S22	$DMSO + OH \longrightarrow 0.95\ CH_3SO_2H + 0.95\ CH_3O_2 + 0.05\ DMSO_2$	2	8.7×10^{-11}		Urbanski et al. (1998)
S23	$CH_3SO_2H + OH \longrightarrow 0.95\ CH_3SO_2 + 0.05\ CH_3SO_3H + 0.05\ HO_2$	2	$9.\times10^{-11}$		Kukui et al. (2003)
S24	$CH_3SO_2H + NO_3 \longrightarrow CH_3SO_2 + HNO_3$	2	1.0×10^{-13}		Yin et al. (1990)
Cl reactions					
Cl1	$Cl + O_3 \longrightarrow ClO + O_2$	2	2.8×10^{-11}	-250	Atkinson et al. (2004)
Cl2	$Cl + HO_2 \longrightarrow HCl + O_2$	2	1.8×10^{-11}	170	Sander et al. (2003)
Cl3	$Cl + HO_2 \longrightarrow ClO + OH$	2	4.1×10^{-11}	-450	Sander et al. (2003)
Cl4	$Cl + H_2O_2 \longrightarrow HCl + HO_2$	2	1.1×10^{-11}	-980	Atkinson et al. (2004)
Cl5	$Cl + CH_3O_2 \longrightarrow 0.5ClO + 0.5HCHO + 0.5HO_2 + 0.5HCl + 0.5CO + 0.5H_2O$	2	1.6×10^{-10}		Sander et al. (2003)
Cl6	$Cl + NO_3 \longrightarrow ClO + NO_2$	2	2.4×10^{-11}		Sander et al. (2003)
Cl7	$Cl + CH_4 \xrightarrow{O_2} HCl + CH_3O_2$	2	9.6×10^{-12}	-1360	Atkinson et al. (2004)
Cl8	$Cl + C_2H_6 \xrightarrow{O_2} HCl + C_2H_5O_2$	2	7.7×10^{-11}	-90	Sander et al. (2003)
Cl9	$Cl + C_2H_4 \xrightarrow{O_2} HCl + C_2H_5O_2$	2	$1.\times10^{-10}$		see note
Cl10	$Cl + HCHO \xrightarrow{O_2} HCl + HO_2 + CO$	2	8.1×10^{-11}	-30	Sander et al. (2003)
Cl11	$Cl + ROOH \longrightarrow CH_3O_2 + HCl$	2	5.7×10^{-11}		Wallington et al. (1990), see note
Cl12	$Cl + OClO \longrightarrow ClO + ClO$	2	3.2×10^{-11}	170	Atkinson et al. (2004)
Cl13	$Cl + ClNO_3 \longrightarrow Cl_2 + NO_3$	2	6.5×10^{-12}	135	Sander et al. (2003)
Cl14	$Cl + PAN \longrightarrow HCl + HCHO + NO_3$	2	1.0×10^{-14}		Sander et al. (2003)
Cl15	$Cl + HNO_3 \longrightarrow HCl + NO_2$	2	1.0×10^{-16}		Sander et al. (2003)
Cl16	$Cl + RONO_2 \longrightarrow HCl + NO_2$	2	7.7×10^{-11}		Michalowski et al. (2000)
Cl17	$ClO + OH \longrightarrow Cl + HO_2$	2	7.4×10^{-12}	270	Sander et al. (2003)
Cl18	$ClO + OH \longrightarrow HCl + O_2$	2	6.0×10^{-13}	230	Sander et al. (2003)
Cl19	$ClO + HO_2 \longrightarrow HOCl + O_2$	2	2.2×10^{-12}	340	Atkinson et al. (2004)
Cl20	$ClO + CH_3O_2 \longrightarrow Cl + HCHO + HO_2$	2	3.3×10^{-12}	-115	Sander et al. (2003)
Cl21	$ClO + NO \longrightarrow Cl + NO_2$	2	6.2×10^{-12}	295	Atkinson et al. (2004)
Cl22	$ClO + NO_2 \xrightarrow{M} ClNO_3$	3	[b]		Atkinson et al. (2004)
Cl23	$ClO + ClO \longrightarrow Cl_2O_2$	2	[b]		Atkinson et al. (2004)
Cl24	$ClO + ClO \longrightarrow Cl_2 + O_2$	2	1.0×10^{-12}	-1590	Atkinson et al. (2004)
Cl25	$ClO + ClO \longrightarrow 2\,Cl + O_2$	2	3.0×10^{-11}	-2450	Atkinson et al. (2004)
Cl26	$ClO + ClO \longrightarrow Cl + OClO$	2	3.5×10^{-13}	-1370	Atkinson et al. (2004)
Cl27	$OClO + OH \longrightarrow HOCl + O_2$	2	4.5×10^{-13}	800	Atkinson et al. (2004)
Cl28	$OClO + NO \longrightarrow ClO + NO_2$	2	1.1×10^{-13}	350	Atkinson et al. (2004)
Cl29	$Cl_2O_2 \longrightarrow ClO + ClO$	1	[b]		Atkinson et al. (2004)
Cl30	$HOCl + OH \longrightarrow ClO + H_2O$	2	3.0×10^{-12}	-500	Sander et al. (2003)
Cl31	$HCl + OH \longrightarrow H_2O + Cl$	2	1.8×10^{-12}	-240	Atkinson et al. (2004)
Cl32	$ClNO_2 + OH \longrightarrow HOCl + NO_2$	2	2.4×10^{-12}	-1250	Atkinson et al. (2004)
Cl33	$ClNO_3 + OH \longrightarrow 0.5\ ClO + 0.5\ HNO_3 + 0.5\ HOCl + 0.5\ NO_3$	2	1.2×10^{-12}	-330	Atkinson et al. (2004)
Cl34	$ClNO_3 \longrightarrow ClO + NO_2$	1	[b]		Anderson and Fahey (1990)
Cl35	$OClO + h\nu \xrightarrow{O_2,\ O_3} O_3 + ClO$	1	[a]		DeMore et al. (1997)
Cl36	$Cl_2O_2 + h\nu \longrightarrow Cl + Cl + O_2$	1	[a]		DeMore et al. (1997)
Cl37	$Cl_2 + h\nu \longrightarrow 2\,Cl$	1	[a]		DeMore et al. (1997)
Cl38	$HOCl + h\nu \longrightarrow Cl + OH$	1	[a]		DeMore et al. (1997)
Cl39	$ClNO_2 + h\nu \longrightarrow Cl + NO_2$	1	[a]		DeMore et al. (1997)
Cl40	$ClNO_3 + h\nu \longrightarrow Cl + NO_3$	1	[a]		DeMore et al. (1997)
Br reactions					
Br1	$Br + O_3 \longrightarrow BrO + O_2$	2	1.7×10^{-11}	-800	Atkinson et al. (2004)
Br2	$Br + HO_2 \longrightarrow HBr + O_2$	2	7.7×10^{-12}	-450	Atkinson et al. (2004)
Br3	$Br + C_2H_4 \xrightarrow{O_2} HBr + C_2H_5O_2$	2	$5.\times10^{-14}$		see note
Br4	$Br + HCHO \xrightarrow{O_2} HBr + CO + HO_2$	2	1.7×10^{-11}	-800	Sander et al. (2003)

n reaction order, [a] photolysis rates calculated online, [b] special rate functions.

Table 2 - Gas phase reactions.

no	reaction	n	A [$(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}$]	$-E_a$ / R [K]	reference
Br5	$\mathrm{Br + ROOH \longrightarrow CH_3O_2 + HBr}$	2	2.66×10^{-12}	-1610	Mallard et al. (1993), see note
Br6	$\mathrm{Br + NO_2 \longrightarrow BrNO_2}$	2	^{b}		Sander et al. (2003)
Br7	$\mathrm{Br + BrNO_3 \longrightarrow Br_2 + NO_3}$	2	4.9×10^{-11}		Orlando and Tyndall (1996)
Br8	$\mathrm{BrO + OH \longrightarrow Br + HO_2}$	2	1.8×10^{-11}	250	Atkinson et al. (2004)
Br9	$\mathrm{BrO + HO_2 \longrightarrow HOBr + O_2}$	2	4.5×10^{-12}	500	Atkinson et al. (2004)
Br10	$\mathrm{BrO + CH_3O_2 \longrightarrow HOBr + HCHO}$	2	$4.1\ \times10^{-12}$		Aranda et al. (1997)
Br11	$\mathrm{BrO + CH_3O_2 \longrightarrow Br + HCHO + HO_2}$	2	$1.6\ \times10^{-12}$		Aranda et al. (1997)
Br12	$\mathrm{BrO + HCHO \xrightarrow{O_2} HOBr + CO + HO_2}$	2	1.5×10^{-14}		Hansen et al. (1999)
Br13	$\mathrm{BrO + NO \longrightarrow Br + NO_2}$	2	8.7×10^{-12}	260	Atkinson et al. (2004)
Br14	$\mathrm{BrO + NO_2 \xrightarrow{M} BrNO_3}$	3	^{b}		Atkinson et al. (2004)
Br15	$\mathrm{BrO + BrO \longrightarrow 2\,Br + O_2}$	2	2.4×10^{-12}	40	Sander et al. (2003)
Br16	$\mathrm{BrO + BrO \longrightarrow Br_2 + O_2}$	2	2.9×10^{-14}	860	Sander et al. (2003)
Br17	$\mathrm{HBr + OH \longrightarrow Br + H_2O}$	2	5.5×10^{-12}	205	Atkinson et al. (2004)
Br18	$\mathrm{BrNO_3 \longrightarrow BrO + NO_2}$	2	^{b}		Orlando and Tyndall (1996)
Br19	$\mathrm{BrO} + h\nu \xrightarrow{\mathrm{O_2}} \mathrm{Br + O_3}$	1	^{a}		DeMore et al. (1997)
Br20	$\mathrm{Br_2} + h\nu \longrightarrow \mathrm{2\,Br}$	1	^{a}		Hubinger and Nee (1995)
Br21	$\mathrm{HOBr} + h\nu \longrightarrow \mathrm{Br + OH}$	1	^{a}		Ingham et al. (1999)
Br22	$\mathrm{BrNO_2} + h\nu \longrightarrow \mathrm{Br + NO_2}$	1	^{a}		Scheffler et al. (1997)
Br23	$\mathrm{BrNO_3} + h\nu \longrightarrow \mathrm{Br + NO_3}$	1	^{a}		DeMore et al. (1997)
Br24	$\mathrm{Br_2 + OH \longrightarrow HOBr + Br}$	2	2.0×10^{-11}	240	Atkinson et al. (2004); Oum et al. (1998)
Br25	$\mathrm{CH_3Br + OH \longrightarrow H_2O + Br}$	2	1.7×10^{-12}	-1215	Atkinson et al. (2003)
Br26	$\mathrm{CHBr_3 + OH \longrightarrow H_2O + Br}$	2	1.35×10^{-12}	-600	Atkinson et al. (2003)
interhalogen reactions					
Hx1	$\mathrm{Cl + BrCl \longrightarrow Br + Cl_2}$	2	1.5×10^{-11}		Mallard et al. (1993)
Hx2	$\mathrm{Cl + Br_2 \longrightarrow BrCl + Br}$	2	1.2×10^{-10}		Mallard et al. (1993)
Hx3	$\mathrm{Br + OClO \longrightarrow BrO + ClO}$	2	2.6×10^{-11}	-1300	Atkinson et al. (2004)
Hx4	$\mathrm{Br + Cl_2 \longrightarrow BrCl + Cl}$	2	1.1×10^{-15}		Mallard et al. (1993)
Hx5	$\mathrm{Br + BrCl \longrightarrow Br_2 + Cl}$	2	3.3×10^{-15}		Mallard et al. (1993)
Hx6	$\mathrm{BrO + ClO \longrightarrow Br + OClO}$	2	1.6×10^{-12}	430	Atkinson et al. (2004)
Hx7	$\mathrm{BrO + ClO \longrightarrow Br + Cl + O_2}$	2	2.9×10^{-12}	220	Atkinson et al. (2004)
Hx8	$\mathrm{BrO + ClO \longrightarrow BrCl + O_2}$	2	5.8×10^{-13}	170	Atkinson et al. (2004)
Hx9	$\mathrm{BrCl} + h\nu \longrightarrow \mathrm{Br + Cl}$	1	^{a}		DeMore et al. (1997)

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions.

Notes: The rates for ROOH were assumed as that of $\mathrm{CH_3OOH}$; $\mathrm{C_2H_4}$ is used as generic alkene as in the Lurmann et al. (1986) mechanism. The rate coefficients are calculated with $k = A \times \exp(\frac{-E_a}{RT})$.

Table 3: Aqueous phase reactions.

no	reaction	n	A [(cm^{-3})$^{1-n}$s^{-1}]	$-E_a$ / R [K]	reference
Ox and HOx reactions					
O1	$O_3 + OH \longrightarrow HO_2$	2	1.1×10^{8}		Sehested et al. (1984)
O2	$O_3 + O_2^- \longrightarrow OH + OH^-$	2	1.5×10^{9}		Sehested et al. (1983)
O3	$OH + OH \longrightarrow H_2O_2$	2	5.5×10^{9}		Buxton et al. (1988)
O4	$OH + HO_2 \longrightarrow H_2O$	2	7.1×10^{9}		Sehested et al. (1968)
O5	$OH + O_2^- \longrightarrow OH^-$	2	1.0×10^{10}		Sehested et al. (1968)
O6	$OH + H_2O_2 \longrightarrow HO_2$	2	2.7×10^{7}	-1684	Christensen et al. (1982)
O7	$HO_2 + HO_2 \longrightarrow H_2O_2$	2	9.7×10^{5}	-2500	Christensen and Sehested (1988)
O8	$HO_2 + O_2^- \xrightarrow{H^+} H_2O_2$	2	1.0×10^{8}	-900	Christensen and Sehested (1988)
O9	$O(^3P) + O_2 \longrightarrow O_3$	2	4.0×10^{9}		Kläning et al. (1984)
NOy reactions					
N1	$HONO + OH \longrightarrow NO_2$	2	1.0×10^{10}		assumed =N7 Barker et al. (1970)
N2	$HONO + H_2O_2 \xrightarrow{H^+} HNO_3$	3	4.6×10^{3}	-6800	Damschen and Martin (1983)
N3	$NO_3 + OH^- \longrightarrow NO_3^- + OH$	2	8.2×10^{7}	-2700	Exner et al. (1992)
N4	$NO_2 + NO_2 \longrightarrow HNO_3 + HONO$	2	1.0×10^{8}		Lee and Schwartz (1981)
N5	$NO_2 + HO_2 \longrightarrow HNO_4$	2	1.8×10^{9}		Warneck (1999)
N6	$NO_2^- + O_3 \longrightarrow NO_3^- + O_2$	2	5.0×10^{5}	-6950	Damschen and Martin (1983)
N7	$NO_2^- + OH \longrightarrow NO_2 + OH^-$	2	1.0×10^{10}		Barker et al. (1970)
N8	$NO_4^- \longrightarrow NO_2^- + O_2$	1	8.0×10^{-1}		Warneck (1999)
N9	$O(^3P) + NO_2^- \longrightarrow NO_3^-$	2	1.48×10^{9}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
N10	$O(^3P) + NO_3^- \longrightarrow NO_2^- + O_2$	2	2.24×10^{8}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
N11	$NO_2 + NO_2 \longrightarrow NO_2^- + NO_3^- + 2\,H^+$	2	1.0×10^{8}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
N12	$NO + NO_2 \longrightarrow NO_2^- + NO_2^- + 2\,H^+$	2	2.0×10^{8}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
N13	$NO + OH \longrightarrow NO_2^- + H^+$	2	2.0×10^{10}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
N14	$NO_2 + OH \longrightarrow NO_3^- + H^+$	2	1.3×10^{9}		Boxe and Saiz-Lopez (2008); Mack and Bolton (1999)
organic reactions					
C1	$HCHO + OH \longrightarrow HCOOH + HO_2$	2	7.7×10^{8}	-1020	Chin and Wine (1994)
C2	$HCOOH + OH \longrightarrow HO_2 + CO_2$	2	1.1×10^{8}	-991	Chin and Wine (1994)
C3	$HCOO^- + OH \longrightarrow OH^- + HO_2 + CO_2$	2	3.1×10^{9}	-1240	Chin and Wine (1994)
C4	$CH_3O_2 + HO_2 \longrightarrow CH_3OOH$	2	4.3×10^{5}		estimated by Jacob (1986)
C5	$CH_3O_2 + O_2^- \longrightarrow CH_3OOH + OH^-$	2	5.0×10^{7}		estimated by Jacob (1986)
C6	$CH_3OH + OH \longrightarrow HCHO + HO_2$	2	9.7×10^{8}		Buxton et al. (1988)
C7	$CH_3OOH + OH \longrightarrow CH_3O_2$	2	2.7×10^{7}	-1715	estimated by Jacob (1986)
C8	$CH_3OOH + OH \longrightarrow HCHO + OH$	2	1.1×10^{7}	-1715	estimated by Jacob (1986)

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions

Table 3 - Aqueous phase reactions.

no	reaction	n	A [$(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}$]	$-E_a$ / R [K]	reference
C9	$CO_3^- + O_2^- \longrightarrow HCO_3^- + OH^-$	2	6.5×10^8		Ross et al. (1992)
C10	$CO_3^- + H_2O_2 \longrightarrow HCO_3^- + HO_2$	2	4.3×10^5		Ross et al. (1992)
C11	$CO_3^- + HCOO^- \longrightarrow HCO_3^- + HCO_3^- + HO_2$	2	1.5×10^5		Ross et al. (1992)
C12	$HCO_3^- + OH \longrightarrow CO_3^-$	2	8.5×10^6		Ross et al. (1992)
C13	$DOM + OH \longrightarrow HO_2$	2	5.0×10^9		estimated by (C. Anastasio, pers. comm.) from Ross et al. (1998)
	S reactions				
S1	$SO_3^- + O_2 \longrightarrow SO_5^-$	2	1.5×10^9		Huie and Neta (1987)
S2	$HSO_3^- + O_3 \longrightarrow SO_4^{2-} + H^+ + O_2$	2	3.7×10^5	-5500	Hoffmann (1986)
S3	$SO_3^{2-} + O_3 \longrightarrow SO_4^{2-} + O_2$	2	1.5×10^9	-5300	Hoffmann (1986)
S4	$HSO_3^- + OH \longrightarrow SO_3^-$	2	4.5×10^9		Buxton et al. (1988)
S5	$SO_3^{2-} + OH \longrightarrow SO_3^- + OH^-$	2	5.5×10^9		Buxton et al. (1988)
S6	$HSO_3^- + HO_2 \longrightarrow SO_4^{2-} + OH + H^+$	2	3.0×10^3		upper limit D. Sedlak pers. comm. with R. Sander
S7	$HSO_3^- + O_2^- \longrightarrow SO_4^{2-} + OH$	2	3.0×10^3		upper limit D. Sedlak pers. comm. with R. Sander
S8	$HSO_3^- + H_2O_2 \longrightarrow SO_4^{2-} + H^+$	2	$5.2\times10^6 \times \frac{[H^+]}{[H^+]+0.1M}$	-3650	Damschen and Martin (1983)
S9	$HSO_3^- + NO_2 \xrightarrow{NO_2} HSO_4^- + 2\,HONO$	2	2.0×10^7		Clifton et al. (1988)
S10	$SO_3^{2-} + NO_2 \xrightarrow{NO_2} SO_4^{2-} + 2\,HONO$	2	2.0×10^7		Clifton et al. (1988)
S11	$HSO_3^- + NO_3 \longrightarrow SO_3^- + NO_3^- + H^+$	2	1.4×10^9	-2000	Exner et al. (1992)
S12	$HSO_3^- + HNO_4 \longrightarrow HSO_4^- + NO_3^- + H^+$	2	3.1×10^5		Warneck (1999)
S13	$HSO_3^- + CH_3OOH \xrightarrow{H^+} SO_4^{2-} + H^+ + CH_3OH$	3	1.6×10^7	-3800	Lind et al. (1987)
S14	$SO_3^{2-} + CH_3OOH \xrightarrow{H^+} SO_4^{2-} + CH_3OH$	3	1.6×10^7	-3800	Lind et al. (1987)
S15	$HSO_3^- + HCHO \longrightarrow CH_2OHSO_3^-$	2	4.3×10^{-1}		Boyce and Hoffmann (1984)
S16	$SO_3^{2-} + HCHO \xrightarrow{H^+} CH_2OHSO_3^-$	2	1.4×10^4		Boyce and Hoffmann (1984)
S17	$CH_2OHSO_3^- + OH^- \longrightarrow SO_3^{2-} + HCHO$	2	3.6×10^3		Seinfeld and Pandis (1998)
S18	$HSO_3^- + HSO_5^- \xrightarrow{H^+} 2\,SO_4^{2-} + 2\,H^+$	3	7.1×10^6		Betterton and Hoffmann (1988)
S19	$SO_4^- + OH \longrightarrow HSO_5^-$	2	1.0×10^9		Jiang et al. (1992)
S20	$SO_4^- + HO_2 \longrightarrow SO_4^{2-} + H^+$	2	3.5×10^9		Jiang et al. (1992)
S21	$SO_4^- + O_2^- \longrightarrow SO_4^{2-}$	2	3.5×10^9		assumed =S20
S22	$SO_4^- + H_2O \longrightarrow SO_4^{2-} + H^+ + OH$	2	1.1×10^1	-1110	Herrmann et al. (1995)
S23	$SO_4^- + H_2O_2 \longrightarrow SO_4^{2-} + H^+ + HO_2$	2	1.2×10^7		Wine et al. (1989)
S24	$SO_4^- + NO_3^- \longrightarrow SO_4^{2-} + NO_3$	2	5.0×10^4		Exner et al. (1992)
S25	$SO_4^- + HSO_3^- \longrightarrow SO_3^- + SO_4^{2-} + H^+$	2	8.0×10^8		Huie and Neta (1987)
S26	$SO_4^- + SO_3^{2-} \longrightarrow SO_3^- + SO_4^{2-}$	2	4.6×10^8		Huie and Neta (1987)
S27	$SO_4^{2-} + NO_3 \longrightarrow NO_3^- + SO_4^-$	2	1.0×10^5		Logager et al. (1993)
S28	$SO_5^- + HSO_3^- \longrightarrow SO_4^- + SO_4^{2-} + H^+$	2	7.5×10^4		Huie and Neta (1987)
S29	$SO_5^- + SO_3^{2-} \longrightarrow SO_4^- + SO_4^{2-}$	2	9.4×10^6		Huie and Neta (1987)
S30	$SO_5^- + HSO_3^- \longrightarrow SO_3^- + HSO_5^-$	2	2.5×10^4		Huie and Neta (1987); Deister and Warneck (1990)
S31	$SO_5^- + SO_3^{2-} \xrightarrow{H^+} SO_3^- + HSO_5^-$	2	3.6×10^6		Huie and Neta (1987); Deister and Warneck (1990)
S32	$SO_5^- + O_2^- \xrightarrow{H^+} HSO_5^- + O_2$	2	2.3×10^8		Buxton et al. (1996)
S33	$SO_5^- + SO_5^- \longrightarrow$ products	2	1.0×10^8		Ross et al. (1992)

n reaction order, [a] photolysis rates calculated online, [b] special rate functions

Table 3 - Aqueous phase reactions.

no	reaction	n	A [$(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}$]	$-E_a$ / R [K]	reference
S34	$\mathrm{DMS} + \mathrm{O_3} \longrightarrow \mathrm{O_2} + \mathrm{DMSO}$	2	8.6×10^{8}	-2600	Gershenzon et al. (2001)
S35	$\mathrm{DMS} + \mathrm{OH} \longrightarrow 0.5\ \mathrm{CH_3SO_3^-} + 0.5\ \mathrm{CH_3O_2} + 0.5\ \mathrm{HSO_4^-} + \mathrm{HCHO} + \mathrm{H^+}$	2	1.9×10^{10}		Ross et al. (1998)
S36	$\mathrm{DMSO} + \mathrm{OH} \xrightarrow{\mathrm{O_2}} \mathrm{CH_3SO_2^-} + \mathrm{CH_3O_2} + \mathrm{H^+}$	2	4.5×10^{9}		Bardouki et al. (2002)
S37	$\mathrm{CH_3SO_2^-} + \mathrm{OH} \xrightarrow{\mathrm{O_2}} \mathrm{CH_3SO_3^-} + \mathrm{H_2O}$	2	1.2×10^{10}		Bardouki et al. (2002)
S38	$\mathrm{CH_3SO_3^-} + \mathrm{OH} \longrightarrow \mathrm{SO_4^{2-}} + \mathrm{H^+} + \mathrm{CH_3O_2}$	2	1.2×10^{7}		Bonsang et al. (1991)
Cl reactions					
Cl1	$\mathrm{Cl} + \mathrm{H_2O_2} \longrightarrow \mathrm{HO_2} + \mathrm{Cl^-} + \mathrm{H^+}$	2	2.0×10^{9}		Yu (2001)
Cl2	$\mathrm{Cl} + \mathrm{H_2O} \longrightarrow \mathrm{H^+} + \mathrm{ClOH^-}$	2	1.8×10^{5}		Yu (2001)
Cl3	$\mathrm{Cl} + \mathrm{NO_3^-} \xrightarrow{\mathrm{H_2O}} \mathrm{NO_3} + \mathrm{Cl^-}$	2	1.0×10^{8}		Buxton et al. (1999b)
Cl4	$\mathrm{Cl} + \mathrm{DOM} \longrightarrow \mathrm{Cl^-} + \mathrm{HO_2}$	2	5.0×10^{9}		estimated (C. Anastasio, pers. comm.) from Ross et al. (1998)
Cl5	$\mathrm{Cl} + \mathrm{SO_4^{2-}} \longrightarrow \mathrm{SO_4^-} + \mathrm{Cl^-}$	2	2.1×10^{8}		Buxton et al. (1999a)
Cl6	$\mathrm{Cl} + \mathrm{Cl} \longrightarrow \mathrm{Cl_2}$	2	8.8×10^{7}		Wu et al. (1980)
Cl7	$\mathrm{Cl^-} + \mathrm{OH} \longrightarrow \mathrm{ClOH^-}$	2	4.2×10^{9}		Yu (2001)
Cl8	$\mathrm{Cl^-} + \mathrm{O_3} \longrightarrow \mathrm{ClO^-} + \mathrm{O_2}$	2	3.0×10^{-3}		Hoigné et al. (1985)
Cl9	$\mathrm{Cl^-} + \mathrm{NO_3} \longrightarrow \mathrm{NO_3^-} + \mathrm{Cl}$	2	9.3×10^{6}	-4330	Exner et al. (1992)
Cl10	$\mathrm{Cl^-} + \mathrm{SO_4^-} \longrightarrow \mathrm{SO_4^{2-}} + \mathrm{Cl}$	2	2.5×10^{8}		Buxton et al. (1999a)
Cl11	$\mathrm{Cl^-} + \mathrm{HSO_5^-} \longrightarrow \mathrm{HOCl} + \mathrm{SO_4^{2-}}$	2	1.8×10^{-3}	-7352	Fortnum et al. (1960)
Cl12	$\mathrm{Cl^-} + \mathrm{HOCl} + \mathrm{H^+} \longrightarrow \mathrm{Cl_2}$	3	2.2×10^{4}	-3508	Ayers et al. (1996)
Cl13	$\mathrm{Cl_2} \longrightarrow \mathrm{Cl^-} + \mathrm{HOCl} + \mathrm{H^+}$	1	2.2×10^{1}	-8012	Ayers et al. (1996)
Cl14	$\mathrm{Cl_2^-} + \mathrm{OH} \longrightarrow \mathrm{HOCl} + \mathrm{Cl^-}$	2	1.0×10^{9}		Ross et al. (1998)
Cl15	$\mathrm{Cl_2^-} + \mathrm{OH^-} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{OH}$	2	4.0×10^{6}		Jacobi (1996)
Cl16	$\mathrm{Cl_2^-} + \mathrm{HO_2} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{H^+} + \mathrm{O_2}$	2	3.1×10^{9}		Yu (2001)
Cl17	$\mathrm{Cl_2^-} + \mathrm{O_2^-} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{O_2}$		6.0×10^{9}		Jacobi (1996)
Cl18	$\mathrm{Cl_2^-} + \mathrm{H_2O_2} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{H^+} + \mathrm{HO_2}$	2	7.0×10^{5}	-3340	Jacobi (1996)
Cl19	$\mathrm{Cl_2^-} + \mathrm{NO_2^-} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{NO_2}$	2	6.0×10^{7}		Jacobi (1996)
Cl20	$\mathrm{Cl_2^-} + \mathrm{CH_3OOH} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{H^+} + \mathrm{CH_3O_2}$	2	7.0×10^{5}	-3340	assumed by Jacobi (1996)
Cl21	$\mathrm{Cl_2^-} + \mathrm{DOM} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{HO_2}$	2	1.0×10^{6}		estimated (C. Anastasio, pers. comm.) from Ross et al. (1998)
Cl22	$\mathrm{Cl_2^-} + \mathrm{HSO_3^-} \longrightarrow \mathrm{SO_3^-} + \mathrm{Cl^-} + \mathrm{Cl^-} + \mathrm{H^+}$	2	4.7×10^{8}	-1082	Shoute et al. (1991)
Cl23	$\mathrm{Cl_2^-} + \mathrm{SO_3^{2-}} \longrightarrow \mathrm{SO_3^-} + \mathrm{Cl^-} + \mathrm{Cl^-}$	2	6.2×10^{7}		Jacobi et al. (1996)
Cl24	$\mathrm{Cl_2^-} + \mathrm{Cl_2^-} \longrightarrow \mathrm{Cl_2} + 2\,\mathrm{Cl^-}$	2	6.2×10^{9}		Yu (2001)
Cl25	$\mathrm{Cl_2^-} + \mathrm{Cl} \longrightarrow \mathrm{Cl^-} + \mathrm{Cl_2}$	2	2.7×10^{9}		Yu (2001)
Cl26	$\mathrm{Cl_2^-} + \mathrm{DMS} \longrightarrow 0.5\mathrm{CH_3SO_3^-} + 0.5\mathrm{CH_3O_2} + 0.5\mathrm{HSO_4^-} + \mathrm{HCHO} + 2\,\mathrm{Cl^-} + 2\,\mathrm{H^+}$	2	3.0×10^{9}		rate from Ross et al. (1998)
Cl27	$\mathrm{ClOH^-} \longrightarrow \mathrm{Cl^-} + \mathrm{OH}$	1	6.0×10^{9}		Yu (2001)
Cl28	$\mathrm{ClOH^-} + \mathrm{H^+} \longrightarrow \mathrm{Cl} + \mathrm{H_2O}$	2	4.0×10^{10}		Yu (2001)
Cl29	$\mathrm{HOCl} + \mathrm{HO_2} \longrightarrow \mathrm{Cl} + \mathrm{O_2}$	2	7.5×10^{6}		assumed = Cl30 Long and Bielski (1980)
Cl30	$\mathrm{HOCl} + \mathrm{O_2^-} \longrightarrow \mathrm{Cl} + \mathrm{OH^-} + \mathrm{O_2}$	2	7.5×10^{6}		Long and Bielski (1980)
Cl31	$\mathrm{HOCl} + \mathrm{SO_3^{2-}} \longrightarrow \mathrm{Cl^-} + \mathrm{HSO_4^-}$	2	7.6×10^{8}		Fogelman et al. (1989)
Cl32	$\mathrm{HOCl} + \mathrm{HSO_3^-} \longrightarrow \mathrm{Cl^-} + \mathrm{HSO_4^-} + \mathrm{H^+}$	2	7.6×10^{8}		assumed = Cl31 Fogelman et al. (1989)
Cl33	$\mathrm{Cl_2} + \mathrm{HO_2} \longrightarrow \mathrm{Cl_2^-} + \mathrm{H^+} + \mathrm{O_2}$	2	1.0×10^{9}		Bjergbakke et al. (1981)
Cl34	$\mathrm{Cl_2} + \mathrm{O_2^-} \longrightarrow \mathrm{Cl_2^-} + \mathrm{O_2}$	2	1.0×10^{9}		assumed = Cl33 Bjergbakke et al. (1981)
Cl35	$\mathrm{Cl^-} + \mathrm{HNO_4} \longrightarrow \mathrm{HOCl} + \mathrm{NO_3^-}$	2	1.4×10^{-2}		Evans et al. (2003)
Br reactions					
Br1	$\mathrm{Br} + \mathrm{OH^-} \longrightarrow \mathrm{BrOH^-}$	2	1.3×10^{10}		Zehavi and Rabani (1972)

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions

Table 3 - Aqueous phase reactions.

no	reaction	n	A $[(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}]$	$-E_a$ / R [K]	reference
Br2	$Br + DOM \longrightarrow Br^- + HO_2$	2	2.0×10^8		estimated (C. Anastasio, pers. comm.) from Ross et al. (1998)
Br3	$Br^- + OH \longrightarrow BrOH^-$	2	1.1×10^{10}		Zehavi and Rabani (1972)
Br4	$Br^- + O_3 \longrightarrow BrO^- + O_2$	2	2.1×10^2	-4450	Haag and Hoigné (1983)
Br5	$Br^- + NO_3 \longrightarrow Br + NO_3^-$	2	3.8×10^9		Zellner et al. 1996 in Herrmann et al. (2000)
Br6	$Br^- + SO_4^- \longrightarrow Br + SO_4^{2-}$	2	2.1×10^9		Jacobi (1996)
Br7	$Br^- + HSO_5^- \longrightarrow HOBr + SO_4^{2-}$	2	1.0	-5338	Fortnum et al. (1960)
Br8	$Br^- + HOBr + H^+ \longrightarrow Br_2$	3	1.6×10^{10}		Liu and Margerum (2001)
Br9	$Br_2 \longrightarrow Br^- + HOBr + H^+$	1	9.7×10^1	7457	Liu and Margerum (2001)
Br10	$Br_2^- + O_2^- \longrightarrow Br^- + Br^-$	2	1.7×10^8		Wagner and Strehlow (1987)
Br11	$Br_2^- + HO_2 \xrightarrow{H^+} Br_2 + H_2O_2$	2	4.4×10^9		Matthew et al. (2003)
Br12	$Br_2^- + H_2O_2 \longrightarrow Br^- + Br^- + H^+ + HO_2$	2	5.0×10^2		Chameides and Stelson (1992)
Br13	$Br_2^- + Br_2^- \longrightarrow Br^- + Br^- + Br_2$	2	1.9×10^9		Ross et al. (1992)
Br14	$Br_2^- + CH_3OOH \longrightarrow Br^- + Br^- + H^+ + CH_3O_2$	2	1.0×10^5		assumed by Jacobi (1996)
Br15	$Br_2^- + DOM \longrightarrow Br^- + Br^- + HO_2$	2	1.0×10^5		estimated (C. Anastasio, pers. comm.) from Ross et al. (1998)
Br16	$Br_2^- + NO_2^- \longrightarrow Br^- + Br^- + NO_2$	2	1.7×10^7	-1720	Shoute et al. (1991)
Br17	$Br_2^- + HSO_3^- \longrightarrow Br^- + Br^- + H^+ + SO_3^-$	2	6.3×10^7	-782	Shoute et al. (1991)
Br18	$Br_2^- + SO_3^{2-} \longrightarrow Br^- + Br^- + SO_3^-$	2	2.2×10^8	-650	Shoute et al. (1991)
Br19	$Br_2^- + DMS \longrightarrow 0.5\ CH_3SO_3^- + 0.5\ CH_3O_2 + 0.5\ HSO_4^- + HCHO + 2\,Br^- + 2\,H^+$	2	3.2×10^9		rate from Ross et al. (1998)
Br20	$BrOH^- \longrightarrow Br^- + OH$	1	3.3×10^7		Zehavi and Rabani (1972)
Br21	$BrOH^- \longrightarrow Br + OH^-$	1	4.2×10^6		Zehavi and Rabani (1972)
Br22	$BrOH^- + H^+ \longrightarrow Br$	2	4.4×10^{10}		Zehavi and Rabani (1972)
Br23	$BrOH^- + Br^- \longrightarrow Br_2^- + OH^-$	2	1.9×10^8		Zehavi and Rabani (1972)
Br24	$BrO^- + SO_3^{2-} \longrightarrow Br^- + SO_4^{2-}$	2	1.0×10^8		Troy and Margerum (1991)
Br25	$HOBr + HO_2 \longrightarrow Br + O_2$	2	1.0×10^9		Herrmann et al. (1999)
Br26	$HOBr + O_2^- \longrightarrow Br + OH^- + O_2$	2	3.5×10^9		Schwarz and Bielski (1986)
Br27	$HOBr + H_2O_2 \longrightarrow Br^- + H^+ + O_2$	2	1.2×10^6		von Gunten and Oliveras (1998)
Br28	$HOBr + SO_3^{2-} \longrightarrow Br^- + HSO_4^-$	2	5.0×10^9		Troy and Margerum (1991)
Br29	$HOBr + HSO_3^- \longrightarrow Br^- + HSO_4^- + H^+$	2	5.0×10^9		assumed = Br28 Troy and Margerum (1991)
Br30	$Br_2 + HO_2 \longrightarrow Br_2^- + H^+ + O_2$	2	1.1×10^8		Ross et al. (1998)
Br31	$Br_2 + O_2^- \longrightarrow Br_2^- + O_2$	2	5.6×10^9		Ross et al. (1998)
Br32	$Br^- + HNO_4 \longrightarrow HOBr + NO_3^-$	2	5.4×10^{-1}		Evans et al. (2003)
Br33	$Br^- + O_3 + H^+ \longrightarrow HOBr + O_2$	2	11.7		Evans et al. (2003)
mixed halide reactions					
Hx1	$Br^- + HOCl + H^+ \longrightarrow BrCl$	3	1.3×10^6		Liu and Margerum (2001)
Hx2	$Cl^- + HOBr + H^+ \longrightarrow BrCl$	3	2.3×10^{10}		Liu and Margerum (2001)

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions

Table 3 - Aqueous phase reactions.

no	reaction	n	A $[(\mathrm{cm}^{-3})^{1-n}\mathrm{s}^{-1}]$	$-E_a$ / R [K]	reference
Hx3	$\mathrm{BrCl} \longrightarrow \mathrm{Cl}^- + \mathrm{HOBr} + \mathrm{H}^+$	1	3.0×10^{6}		Liu and Margerum (2001)
Hx4	$\mathrm{Br}^- + \mathrm{ClO}^- + \mathrm{H}^+ \longrightarrow \mathrm{BrCl} + \mathrm{OH}^-$	3	3.7×10^{10}		Kumar and Margerum (1987)
Hx5	$\mathrm{Cl_2} + \mathrm{Br}^- \longrightarrow \mathrm{BrCl_2^-}$	2	7.7×10^{9}		Liu and Margerum (2001)
Hx6	$\mathrm{BrCl_2^-} \longrightarrow \mathrm{Cl_2} + \mathrm{Br}^-$	1	1.83×10^{3}		Liu and Margerum (2001)
photolysis					
hv1	$\mathrm{O_3} + h\nu \longrightarrow \mathrm{OH} + \mathrm{OH} + \mathrm{O_2}$		^{a}		equal to gas phase
hv2	$\mathrm{H_2O_2} + h\nu \longrightarrow \mathrm{OH} + \mathrm{OH}$		^{a}		equal to gas phase
hv3	$\mathrm{NO_3^-} + h\nu \xrightarrow{\mathrm{H}^+} \mathrm{NO_2} + \mathrm{OH}$		^{a}		Warneck and Wurzinger (1988); Zellner et al. (1990)
hv4	$\mathrm{NO_2^-} + h\nu \xrightarrow{\mathrm{H}^+} \mathrm{NO} + \mathrm{OH}$		^{a}		Warneck and Wurzinger (1988); Zellner et al. (1990)
hv5	$\mathrm{HOCl} + h\nu \longrightarrow \mathrm{OH} + \mathrm{Cl}$		^{a}		equal to gas phase
hv6	$\mathrm{Cl_2} + h\nu \longrightarrow \mathrm{Cl} + \mathrm{Cl}$		^{a}		equal to gas phase
hv7	$\mathrm{HOBr} + h\nu \longrightarrow \mathrm{OH} + \mathrm{Br}$		^{a}		equal to gas phase
hv8	$\mathrm{Br_2} + h\nu \longrightarrow \mathrm{Br} + \mathrm{Br}$		^{a}		equal to gas phase
hv9	$\mathrm{BrCl} + h\nu \longrightarrow \mathrm{Cl} + \mathrm{Br}$		^{a}		equal to gas phase
hv10	$\mathrm{NO_3^-} + h\nu \longrightarrow \mathrm{NO_2^-} + \mathrm{O(^3P)}$		^{a}		Warneck and Wurzinger (1988); Zellner et al. (1990)
hv11	$\mathrm{O_3} + h\nu \longrightarrow \mathrm{O_2} + \mathrm{O(^3P)}$		^{a}		equal to gas phase

n reaction order, ^{a} photolysis rates calculated online, ^{b} special rate functions

Note: The rate coefficients are calculated with $k = A \times \exp(\frac{-E_a}{RT})$.

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