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Mechanistic investigations of the formation of multifunctional products from the multi-generation ·OH oxidation of styrene

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Abstract. Styrene is a highly reactive aromatic hydrocarbon that has been identified as a key secondary organic aerosol (SOA) precursor. Recent laboratory chamber experiments have identified C_7 and C_8 series compounds as the main components of SOA in the photooxidation of styrene. However, their molecular structures and formation pathways remain largely uncharacterized. Herein, the formation mechanisms of multifunctional products from the multi-generation •OH oxidation of styrene are studied using the quantum chemistry methods. The calculations show that the first generation RO_2 radicals can either proceed unimolecular decomposition to yield benzaldehyde (C_7H_6O), or undergo bimolecular reactions with HO_2•/NO to form the first generation closed-shell C_7- and C_8-products, hydroperoxide 1st-ROOH ($C_8H_{10}O_3$), benzaldehyde, and organic nitrate 1st-$RONO_2$ ($C_8H_9NO_3$). For the second generation •OH oxidation, OH-addition reaction occurring at the *ortho*-site of 1st-ROOH and 1st-$RONO_2$ has a significant dominance. The *ortho*-OH-addition products can proceed through two O_2-addition steps and a cyclization process to produce the peroxide bicyclic peroxy radicals (BPR). BPR can further react with HO_2•/NO to form the second generation closed-shell C_8-products, hydroperoxide 2nd-ROOH ($C_8H_{12}O_8$), organic nitrate 2nd-$RONO_2$ ($C_8H_{10}N_2O_{10}$), and other multifunctional products, in which the first two products have fractional yields of 41.4 % and 4.8 %, respectively. For the third generation •OH oxidation, OH-addition occurring at the C=C double bond of 2nd-ROOH and 2nd-$RONO_2$ has the lowest barrier. The major third generation closed-shell C_8-products are the multifunctional hydroperoxides and organic nitrates. These findings carry important implications for advancing our understanding of the chemical composition and formation mechanisms of aromatic SOA.

1 Introduction

Aromatic compounds are recognized as the significant secondary organic aerosol (SOA) precursors, accounting for 20 %–30 % of the total volatile organic compounds (VOCs) and up to ∼ 60 % of the urban atmosphere (Xu et al., 2020; Yan et al., 2019; Yu et al., 2022a, b; Cabrera-Perez et al., 2016; Iyer et al., 2023; Wang et al., 2017; Bloss et al., 2005; Forstner et al., 1997). The primary sources include the incomplete combustion, solvent evaporation, and industrial emission, and the secondary sources involve the biofuel and biomass burning (Xu et al., 2020; Cabrera-Perez et al., 2016; Li et al., 2019). The most abundant aromatic compounds, including benzene, toluene, ethylbenzene, xylenes, styrene and trimethylbenzenes, are highly present in urban environments (Cabrera-Perez et al., 2016; Koppmann, 2008). The degra-

dation of aromatic compounds initiated by the atmospheric oxidants (e.g., OH radicals, NO_3 radicals, O_3, and Cl atom) leads to the production of multifunctional molecules (e.g., nitroaromatics, dicarbonyls, cresols, epoxides) (Ji et al., 2017; Wu et al., 2014; Fu et al., 2023; Wang and Li, 2021; Wang et al., 2013, 2020b; Zaytsev et al., 2019), significantly contributing to new particle formation (NPF) and SOA formation (up to 50 % in eastern China) in the atmosphere (Wang et al., 2017, 2020a; Garmash et al., 2020; Molteni et al., 2018; Nie et al., 2022).

The secondary organic aerosol formation potential (SOAP) of aromatics is significantly greater than that of alkanes and alkenes during haze episodes in Beijing (Sun et al., 2016). Among these precursors, toluene is the predominant SOA-forming species, contributing more than 16 % of the total SOA, followed by styrene (15 %) and ethylbenzene (9.5 %) (Sun et al., 2016). Styrene is primarily emitted from the anthropogenic activities such as solvent usage and vehicle exhaust (Cho et al., 2014; Wu et al., 2021), which is detected at the ppb levels in urban environments, with the mixing ratios of 0.06–4.50 ppb (Cho et al., 2014; Huang et al., 2019). Styrene has been classified as a hazardous air pollutant in the 1990 Clean Air Act due to the potential mutagen and carcinogen (Environmental Protection Agency (EPA), 1990). Therefore, it is very necessary to investigate the degradation mechanisms of styrene under atmospheric conditions. In general, the atmospheric oxidation of styrene initiated by OH radicals is anticipated to be the dominant daytime sink, and the lifetime is estimated to be ~ 8 h under the conditions of typical OH radicals concentrations ($[\bullet OH] = \sim 2 \times 10^6$ molec. cm^{-3}) (Wu et al., 2021; Shen et al., 2022). Due to the existence of highly reactive vinyl and aromatic groups, OH-initiated oxidation of styrene mainly comprise two kinds of pathways: H-abstraction and OH-addition, in which C_β-site OH-addition reaction is expected to be the predominant pathway (Wu et al., 2021; Wang et al., 2015; Zhang et al., 2024). The formed products can combine with an O_2 molecule leading to the first generation peroxyl radicals, which can further react with NO resulting in the formation of benzaldehyde and formaldehyde. The barrier of the rate-limiting step is predicted to be 28.4 kcal mol^{-1} (Wang et al., 2015), implying that benzaldehyde is unlikely to be the sole primary product in the oxidation of styrene due to their higher barriers. Additionally, carbonyl oxides, formed in the ozonolysis of styrene, serve as the chain units participating in the formation of oligomers (Yu et al., 2022a). The volatility of oligomers decreases dramatically as the successive addition of carbonyl oxides increases, eventually transforming into extremely low volatility organic compounds (ELVOC) and directly participating in NPF.

Experimentally, Cho et al., investigated the kinetics of the reaction styrene $+\bullet$OH at 240–340 K and 1–3 Torr using the mass spectrometry technique (Cho et al., 2014). They found that the addition of OH radicals to the vinyl carbons is dominant, and the determined rate coefficient is $(5.80 \pm 0.49) \times 10^{-11}$ cm^3 molec.$^{-1}$ s^{-1} at room temperature. In the smog chamber experiments, Tajuelo et al. (2019a, b, c) found that the SOA yields from the photolysis and photooxidation of styrene and its homologous species increase with the concentration of initial reactants increasing, and benzaldehyde, benzoyl chloride, acetophenone and formaldehyde are expected to be the primary gas phase products. Yu et al. (2022b) investigated the formation of SOA from styrene in an indoor chamber under different NO_x and RH conditions, and found the SOA yields decrease with increasing RH in both the H_2O_2 and NO_x systems. The C_7 and C_8 species are the main products in the H_2O_2 system, while organic nitrates are the major components in the NO_x system. Although the possible molecular formula and chemical composition of the oxidation products from the reaction styrene $+\bullet$OH are given in the aforementioned studies, the specific molecular structures and formation pathways remain ambiguous. Additionally, to the best of our knowledge, the majority of studies mainly focus on the first generation •OH oxidation products to date, while the formation mechanisms of multifunctional products from the multi-generation •OH oxidation of styrene are still limited.

In the present study, the multi-generation •OH oxidation mechanisms of styrene in the presence of $HO_2\bullet$/NO are investigated using the quantum chemistry methods. The calculated results arising from the first generation •OH oxidation reactions are presented herein for comparison with the available literatures to ascertain the reliability of the employed theoretical method. For the multi-generation •OH oxidation reactions of styrene, all the possible pathways, including H-abstraction, OH-addition, O_2-addition, cyclization, ring-opening, intramolecular H-shifts, C−C bond and O−O bond scission, and HO_2-elimination, are taken into account. Additionally, the saturated concentrations of the formed multifunctional products are estimated to identify the volatility classes.

2 Computational methods

2.1 Electronic structures and energy calculations

The electronic structures and energy calculations of all stationary points, including reactants (R), intermediates (IM), transition states (TS) and products (P), are performed using the Gaussian 16 program (Frisch et al., 2016). Geometric optimizations of all stationary points on the potential energy surfaces (PESs) are carried out at the M06-2X/6-31+g(d,p) level of theory, since it has reliable performance for describing the noncovalent interactions, thermochemical, and kinetics (Zhao and Truhlar, 2008). Harmonic vibrational frequencies are determined at the M06-2X/6-31+g(d,p) theoretical level to confirm the characteristics of all stationary points (a local minimum or a saddle point). The zero-point vibrational energy (ZPVE) is scaled by a factor of 0.967 (Alecu et al., 2010). Intrinsic reaction coordinate (IRC) calcula-

tions are carried out to ascertain the connection of the given TS between the designated local minima R and P (Fukui, 1981). Single point energy calculations are performed at the M06-2X/6-311++G(3df,3pd) level based on the M06-2X/6-31+g(d,p) optimized geometries.

In order to further evaluate the reliability of the computational method employed herein, the single point energies of all the stationary points involved in the initial addition of OH radicals to styrene and intramolecular H-shift reactions of the first generation peroxyl radicals S2-1-x are recalculated using the DLPNO-CCSD(T)/aug-cc-pVTZ method performed using the Orca 6.1 program (Neese, 2025). As shown in Table S1 in the Supplement, the ΔE_a values obtained using the M06-2X/6-311++G(3df,3pd) method are consistent with those derived from the DLPNO-CCSD(T)/aug-cc-pVTZ method. The largest deviation and the average absolute deviation are 1.2 and 0.6 kcal mol^{-1}, respectively, indicating that the computational method employed in this study is reliable. Considering the computational cost, the M06-2X/6-311++G(3df,3pd) method is employed to investigate the formation mechanism of multifunctional products from the multi-generation •OH oxidation of styrene. The energy barrier (ΔE_a) and reaction energy (ΔE_r) are defined as the difference in energy between TS and IM, as well as between P and R.

2.2 Conformer research

RO_2 radicals formed from the addition of O_2 to the carbon-centered site of alkyl radicals R have multiple possible conformers due to the different orientations of O_2 attack (Chen et al., 2021; Fu et al., 2020; Møller et al., 2016, 2020). An initial structure of RO_2 radicals is optimized at the B3LYP/6-31+G(d) level and subsequently used as the starting geometry to perform the conformer search conducted using the Molclus program (Lu, 2024). The resulting structures are initially optimized at the B3LYP/6-31+G(d) level, as this method accurately predicts the relative energy ordering of different conformers (Møller et al., 2016, 2020). For the intramolecular H-shift reactions of RO_2 radicals, the lengths of the O−O, C−H and O−H bonds in the conformational sampling of TSs are constrained to retain the cyclic TS structure. All unique conformers of R, TS and P within 5.0 kcal mol^{-1} with respect to the lowest energy conformer are further optimized at the M06-2X/6-31+g(d,p) level of theory. Then, the single point energy calculations are performed at the M06-2X/6-311++G(3df,3pd) level of theory. RO radicals formed by the bimolecular reactions of RO_2 radicals with HO_2 radicals and NO also have multiple conformers. In order to obtain the lowest energy conformer, a similar methodology is employed in the present study.

2.3 Kinetics calculations

The rate coefficients of unimolecular reactions, including intramolecular H-shifts, cyclization, HO_2-elimination, and C−C bond and C−O bond scissions, are calculated using the RRKM theory along with energy-grained master equation (RRKM-ME) (Holbrook et al., 1996). The rate coefficients of bimolecular reactions, involving H-abstraction and OH-addition, are determined using the traditional transition state theory (TST) (Fernández-Ramos et al., 2007). An asymmetric one-dimensional Eckart model (Eckart, 1930) is employed to consider the tunneling correction factors in the rate coefficient calculations based on RRKM-ME and TST. A single exponential down model in the RRKM-ME calculations is utilized to approximate the collision transfer ($\langle \Delta E \rangle_{\text{down}} = 200\ \text{cm}^{-1}$). The Lennard–Jones parameters of all intermediate species are estimated using the empirical formula as proposed by Gilbert and Smith (1990).

For the intramolecular H-shifts of RO_2 and RO radicals, the rate coefficients are computed using the multiconformer transition state theory (MC-TST) (Møller et al., 2016), which is expressed as Eq. (1) (Møller et al., 2016, 2020; Pasik et al., 2024):

$$k_{\text{MC-TST}} = \kappa \frac{k_B T}{h} \frac{\sum_i^{\text{TS conf.}} \exp\left(\frac{-\Delta E_i}{k_B T}\right) Q_{\text{TS},i}}{\sum_j^{\text{R conf.}} \exp\left(\frac{-\Delta E_j}{k_B T}\right) Q_{\text{R},j}} \exp\left(-\frac{E_{\text{TS}} - E_{\text{R}}}{k_B T}\right) \quad (1)$$

where κ is the Eckart tunneling coefficient, h is Planck's constant, k_B is Boltzmann's constant, and T is the absolute temperature (298.15 K). $Q_{\text{TS},i}$ and $Q_{\text{R},j}$ refer to the partition functions of the corresponding transition state i and reactant j conformers, respectively. ΔE_i and ΔE_j represent the relative electronic energies between the corresponding transition state i and reactant j conformers and the lowest energy conformers, respectively. E_{TS} and E_{R} stand for the electronic energies of the lowest energy transition state and reactant conformers, respectively. All kinetics calculations are carried out using the KiSThelP 2021 and MESMER 6.0 programs (Glowacki et al., 2012; Canneaux et al., 2013).

3 Results and discussion

3.1 First generation •OH oxidation mechanisms of styrene

Styrene is composed of a benzene ring and a vinyl group, and its oxidation initiated by OH radicals may proceed either on the vinyl group or on the benzene ring. Previous literature has demonstrated that the addition of OH radicals to terminal carbon (C_β-site) of a vinyl group in styrene is the dominant pathway, with the branching ratio of 88.2 % (Wu et al., 2021). Therefore, the C_β-site OH-addition reaction is mainly considered in the present study. Figure 1 depicts that this reaction starts with the formation of a pre-reactive complex IM1, and then transforms into an alkyl radical S1-1 via transition state TS1 with a ΔE_a of 0.8 kcal mol^{-1}.

The rate coefficient of C_β-site OH-addition reaction is estimated to be 1.5×10^{-11} cm^3 molec.$^{-1}$ s^{-1} at ambient temperature, which is approximately consistent with the experimental ($1.2–6.2 \times 10^{-11}$ cm^3 molec.$^{-1}$ s^{-1}) and theoretical values ($1.7–2.0 \times 10^{-11}$ cm^3 molec.$^{-1}$ s^{-1}) for the total rate coefficient of the reaction styrene +•OH (Wu et al., 2021; Zhang et al., 2024).

Due to the existence of resonance structures with radical character on the aromatic ring, the resulting S1-1 can readily isomerize into three other species, namely, S1-2, S1-3 and S1-4. The attack of an O_2 molecule on the C-center site of S1-1 leads to the formation of the first generation peroxy radicals S2-1-x ($\Delta E_r > -59.6$ kcal mol^{-1}). The formed S2-1-x includes eight energetically similar conformers due to the different orientations of O_2 attack. In order to distinguish the different conformers, the subscript letter x is used in the present study. The energy ordering of different conformers follows an alphabetical sequence, in which letter a denotes the lowest energy conformer. The Boltzmann population of different conformers in S2-1-x is listed in Table S2 in the Supplement.

For the unimolecular decomposition reactions of S2-1-x, there are three kinds of pathways. One is the intramolecular H-shift reactions, where the hydrogen atom migrates from the $-CH_2$, $-CH$ and $-OH$ groups to the terminal oxygen atom of the $-OO$ group leading to various alkyl and alkoxyl radicals. Among these competing H-shift reactions, the hydrogen atom at the $-OH$ group can be transferred via a six-membered ring transition state (1,5-H shift) to yield an alkoxyl radical S3-1-a, which exhibits the lowest barrier ($\Delta E_a = 21.0$ kcal mol^{-1}). The resulting S3-1-a can undergo the $C_\alpha-C_\beta$ bond cleavage to produce a formaldehyde and an alkyl radical S4-1-a ($\Delta E_a = 0.8$ kcal mol^{-1}), followed by an OH radical release to form benzaldehyde ($\Delta E_a = 0.1$ kcal mol^{-1}). The rate coefficients for the aforementioned three pathways are calculated to be 2.7×10^{-4}, 4.6×10^{10} and 7.2×10^{10} s^{-1}, respectively. Based on the values of ΔE_a and the corresponding rate coefficients, it can be concluded that the 1,5-H shift reaction is the rate-determining step in the formation of benzaldehyde. The other is the cyclization, where the $-OO$ group attacks the C=C double bond in the benzene ring forming a cyclic peroxide alkyl radical S3-1-c ($\Delta E_a = 33.6$ kcal mol^{-1}). The last is the HO_2-elimination, where a concerted process of $C_\alpha-O$ and $C_\beta-H$ bonds scission forms a closed-shell species S3-1-b and a HO_2 radical byproduct ($\Delta E_a = 33.3$ kcal mol^{-1}). The aforementioned results show that the cyclization and HO_2-eliminaiton reactions are less importance due to their higher barriers.

As depicted in Fig. S1 in the Supplement, the formations of the first generation peroxy radicals S2-2-x from the association reaction S1-2 + O_2 are strongly endothermic ($\Delta E_r = 8.1-10.4$ kcal mol^{-1}), suggesting that they have a significant potential to redissociate back to reactants S1-2 and O_2. The resulting S2-2-x can undergo through various intramolecular H-shifts to yield distinct C-centered and O-centered radicals. Among these competing H-shift pathways, hydrogen transfer from the $-OH$ group to the terminal oxygen of $-OO$ group has the lowest barrier ($\Delta E_a = 17.4$ kcal mol^{-1}). A similar conclusion is also obtained from the association reactions S1-3 + O_2 ($\Delta E_r = 6.6-7.1$ kcal mol^{-1}) and S1-4 + O_2 ($\Delta E_r = 8.1-11.1$ kcal mol^{-1}) that the formations of the first generation peroxyl radicals S2-3-x and S2-4-x are thermochemically unfavorable, and their subsequent intramolecular H-shift barriers are considerably high (Figs. S2 and S3 in the Supplement). Therefore, in the present study, we mainly focus on the subsequent reaction mechanisms of S2-1-x under both low and high NO_x conditions.

In the low-NO_x conditions, the bimolecular reaction with HO_2 radicals is expected to be the dominant sink for RO_2 radicals (Orlando and Tyndall, 2012; Vereecken et al., 2015). Previous studies have reported that the rate coefficient $k_{RO_2+HO_2}$ for the reactions of alkyl peroxyl radicals with HO_2 radicals is 1.7×10^{-11} cm^3 molec.$^{-1}$ s^{-1} (Atkinson and Arey, 2003; Boyd et al., 2003). The typical atmospheric concentration of HO_2 radicals is 20–40 pptv (Wang et al., 2017; Bianchi et al., 2019), resulting in the pseudo-first-order rate constant $k'_{RO_2+HO_2} = k_{RO_2+HO_2}$ [HO_2] of 0.01–0.02 s^{-1}. The isomerization reaction of RO_2 radicals is competitive with the bimolecular reactions with HO_2 radicals only when the rate coefficient of intramolecular H-shifts exceeds 0.01–0.02 s^{-1}. In the high-NO_x conditions, the bimolecular reaction of RO_2 radicals with NO is considered to be a dominant sink (Orlando and Tyndall, 2012; Vereecken et al., 2015). The rate coefficient k_{RO_2+NO} for the reaction of alkyl peroxyl radicals with NO is determined to be 9.0×10^{-12} cm^3 molec.$^{-1}$ s^{-1} (Atkinson and Arey, 2003; Bianchi et al., 2019). The typical atmospheric concentration of NO is 0.4–40 ppbv (Wang et al., 2017; Bianchi et al., 2019), leading to the pseudo-first-order rate constant $k'_{RO_2+NO} = k_{RO_2+NO}$ [NO] of 0.1–10 s^{-1}. The intramolecular H-shift reaction of RO_2 radicals can compete with the bimolecular reaction with NO only when the rate coefficient of the former case exceeds 10 s^{-1}. Therefore, we use the $k'_{RO_2+HO_2}$ (0.01–0.02 s^{-1}) and k'_{RO_2+NO} (0.1–10 s^{-1}) values as thresholds to evaluate the relative importance of the isomerization reactions of RO_2 radicals under both low- and high-NO_x conditions. Previous studies have also employed the same methodology to evaluate the relative importance of isomerization and bimolecular reactions of RO_2 radicals during the OH-initiated oxidation of organophosphate esters and alkylbenzenes (Wang et al., 2017; Fu et al., 2024). For the intramolecular H-shift reactions of S2-1-x, the rate coefficient $k_{MC\text{-}TST}$ is estimated to be 1.6×10^{-4} s^{-1}, which is 2–4 orders of magnitude lower than $k'_{RO_2+HO_2}$ and k'_{RO_2+NO}, indicating that the isomerization reaction of S2-1-x is less competitive than the bimolecular reactions with HO_2 radicals and NO.

In the presence of NO, the bimolecular reactions of S2-1-x with NO initially proceed via oxygen-to-oxygen coupling to

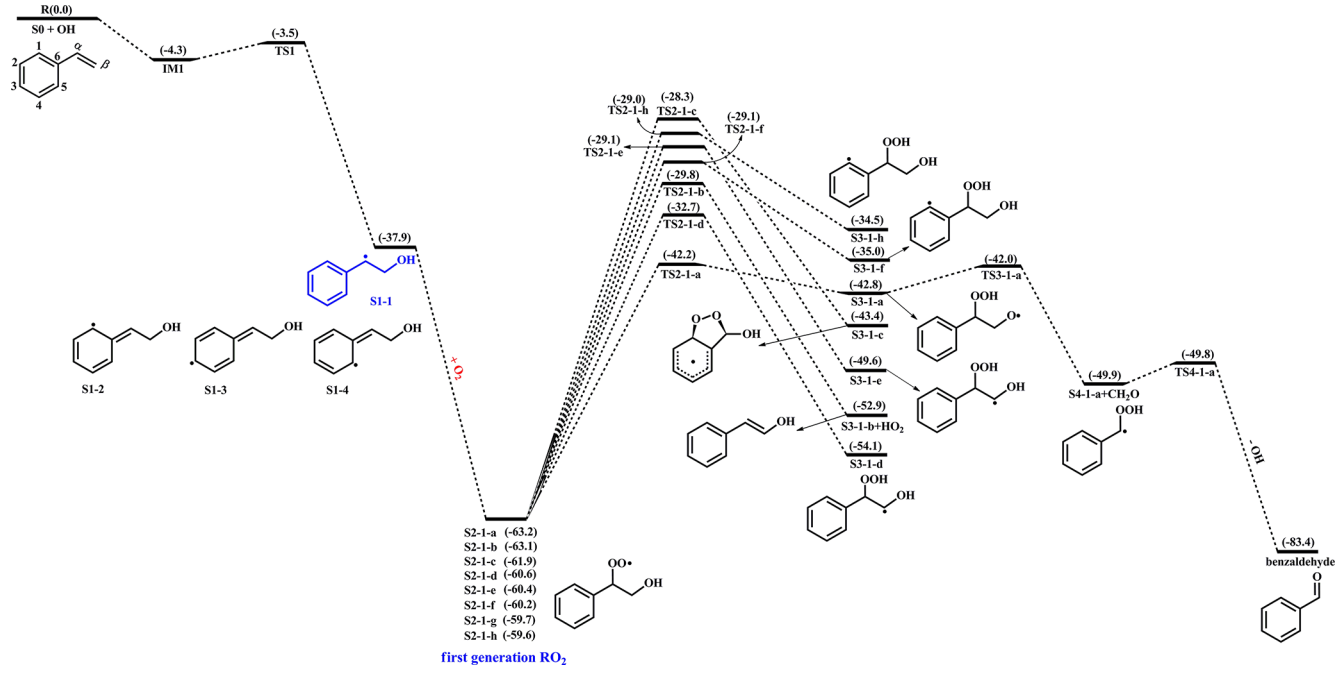

Figure 1. PES for the first-stage oxidation of styrene initiated by OH radicals and the isomerization reactions of S2-1-x at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

yield organic nitrites ROONO, which subsequently decompose into benzaldehyde and CH_2OH radical or isomerize to organic nitrates $RONO_2$. The energy barrier of the rate-limiting step predicted in Wang's study for the formation of benzaldehyde is 28.4 kcal mol^{-1}, which is approximately 4.0 kcal mol^{-1} greater than that for the formation of $RONO_2$ (Wang et al., 2015). In the absence of NO, the hydroperoxides ROOH formed from the bimolecular reaction of S2-1-x with HO_2 radicals are anticipated to be the dominate products. The aforementioned results are further confirmed by the recent smog chamber experiment study that C_7 and C_8 series products, as well as organic nitrates are the main components of SOA in the OH-initiated oxidation of styrene under different NO_x conditions (Yu et al., 2022b). Considering that the extensive studies on the OH-initiated oxidation of benzaldehyde have done (Sebbar et al., 2011; Zhao et al., 2022; Iuga et al., 2008), this study primarily focuses on the multi-generation •OH oxidation mechanisms of ROOH and $RONO_2$ under both low- and high-NO_x conditions.

3.2 Second generation •OH oxidation mechanisms of 1st-ROOH and 1st-$RONO_2$

The first generation products, including hydroperoxides 1st-ROOH and organic nitrates 1st-$RONO_2$, include multiple conformers. To obtain the global minimum of 1st-ROOH and 1st-$RONO_2$, the conformer search is performed by using the Molclus program. The resulting structures are initially optimized at the M06-2X/6-31+g(d,p) level, then the single point energies are calculated at the M06-2X/6-311++G(3df,3pd) level. The global minimum structures of 1st-ROOH (S4) and 1st-$RONO_2$ (S5) are presented in Fig. S4 in the Supplement.

3.2.1 The oxidation mechanism of 1st-ROOH initiated by OH radicals

The reaction 1st-ROOH (S4) +•OH proceeds through the addition of OH radicals to either side of the benzene ring to yield various alkyl radicals, as depicted in Fig. 2. In the present study, *syn*-OH-addition is defined as the scenario in which the addition of OH radicals occurs at the same side as the −OOH group, while *anti*-OH-addition is referred to the scenario in which the addition of OH radicals occurs at the opposite side as the −OOH group. For the *syn*-OH-addition reactions, the addition of OH radicals to the C1-site of 1st-ROOH (S4) exhibits the lowest barrier ($\Delta E_a =$ 3.6 kcal mol^{-1}) due to the stability of the formed product, $P_{S4\text{-}add1'}$. A similar conclusion is also obtained from the *anti*-OH-addition reactions that the OH-addition pathway occurring at the C1-site is favorable ($\Delta E_a = 0.8$ kcal mol^{-1}). Notably, the preferred OH-addition pathway in the *anti*-OH-addition reactions exhibits greater competitiveness compared to that in the *syn*-OH-addition reactions. It can be explained by the greater steric hindrance present in the latter reaction. In order to further evaluate the reliability of our results, ΔE_a of all the *syn*-OH-addition and *anti*-OH-addition reactions are recalculated using the DLPNO-CCSD(T)/aug-cc-pVTZ//M06-2X/6-311+G(d,p) method. As shown in Ta-

ble S3 in the Supplement, the ΔE_a values obtained using the M06-2X/6-311++G(3df,3pd) method are in good agreement with those derived from the DLPNO-CCSD(T)/aug-cc-pVTZ method. The largest deviation and the average absolute deviation are 1.2 and 0.9 kcal mol^{-1}, respectively, indicating that the M06-2X/6-311++G(3df,3pd) method employed in this study is reliable. Based on the values of ΔE_a obtained using the DLPNO-CCSD(T)/aug-cc-pVTZ method, it can also be concluded that the addition of OH radicals to C1-site, occurring at the opposite direction relative to the −OOH group, is energetically favorable. The rate coefficients of the addition of OH radicals to the different sites of 1st-ROOH are calculated to be 8.2×10^{-12} (C1-site), 5.8×10^{-15} (C2-site), 8.3×10^{-15} (C3-site), 8.6×10^{-15} (C4-site), 2.7×10^{-12} (C5-site) and 4.1×10^{-13} (C6-site) cm^3 molec.$^{-1}$ s^{-1}, respectively. The branching ratios for •OH addition to the C1, C5 and C6 sites are predicted to be 72.4 %, 23.8 % and 3.6 %, respectively, while the sum of branching ratios for •OH addition to other carbon sites is less than 1 %.

Our result is opposite to Zhang's finding that the addition of OH radicals to C6-site would be the most favorable pathway (Zhang et al., 2024). The discrepancy can be explained by the following three factors: (1) The 1st-ROOH conformer selected in the Zhang's study is not the global minimum. In the present study, the global minimum conformer of 1st-ROOH, identified through the conformer search, is found to be 2.2 kcal mol^{-1} lower than the 1st-ROOH structure selected in the Zhang's study. (2) The pre-reactive complexes are not considered in the Zhang's study. The addition of OH radicals to C1-, C2-, C3- and C6-sites, occurring at the opposite direction relative to the −OOH group, are merely considered in the Zhang's study. They found that the apparent energy barrier of the addition of OH radicals to C6-site is smallest, and is therefore expected to be the favorable pathway. Actually, these OH-addition reactions are modulated by the pre-reactive complexes. It may be inappropriate to determine the favorable pathway based solely on apparent activation energy without considering the pre-reaction complexes. (3) From a geometric perspective, the addition of OH radicals to C6-site is associated with greater steric hindrance compared to other sites, as C6-atom connects with a larger functional group. Base on the aforementioned discussions, we believe that the addition of OH radicals to C6-site is unlikely to be the dominant pathway. Our calculations also confirm that the addition of OH radicals to C6-site is less importance compared to that at the C1-site.

Our conclusion is further supported by the reaction toluene +•OH that the *ortho*-OH-addition reaction exhibits significant dominance, with the branching ratio of up to 69.8 %–75.8 % (Ji et al., 2017; Zhang, 2019; Wu et al., 2020). Considering the high reactivity of *ortho*-OH-addition in the reactions toluene +•OH and 1st-ROOH (S4) +•OH, the substitute effects of the $-CH_3$ and −OOH groups are explicitly discussed in the present study. Notably, the $-CH_3$ group in toluene is bonded to the C6 atom, and the −OOH group in 1st-ROOH is bonded to the Cα atom, as depicted in Fig. S5 in the Supplement. The optimized geometries of toluene and 1st-ROOH and the NPA atomic charges of all the carbon atoms in the benzene ring are displayed in Fig. S5. The C−C bond lengths and the C−C−C bond angles in the benzene ring of toluene are approximately 1.39 Å and 120°, respectively, which are consistent with those in the benzene ring of 1st-ROOH. The aforementioned results show the effect of the $-CH_3$ and −OOH groups on the geometric structure of benzene ring is negligible. From the perspective of NPA atomic charges, the charges on the C1 (−0.246 e) and C5 (−0.246 e) atoms are more greater than those on the other carbon atoms in the benzene ring of toluene. And the OH-adduct formed from the *ortho*-OH-addition reaction exhibits the greater stability. These results indicate that the $-CH_3$ group is a typical *ortho*-directing substituent and exerts an activating effect on the *ortho*-site of the benzene ring, which explains why the *ortho*-OH-addition reaction is predominant in the reaction toluene +•OH. Compared with the charges on the carbon atoms in the benzene ring of toluene, the charges on C1 and C6 atoms increase by 0.013 e and 0.057 e, respectively, in 1st-ROOH, which can be attributed to the electron-withdrawing effect of the −OOH group. The charge on the C1 atom (−0.259 e) is the highest, and the stability of the resulting OH-adduct is the greatest, implying that the addition of OH radicals to C1-site is dominant in the reaction 1st-ROOH +•OH. Therefore, a direct comparison of the favorable OH-addition pathway in the reactions toluene +•OH and 1st-ROOH (S4) +•OH is performed in this study.

The formed product $P_{S4\text{-}add1}$ includes two conjugate double bonds (C_2=C_3 and C_4=C_5), which can readily isomerize to $P_{S4\text{-}add2}$ and $P_{S4\text{-}add3}$, as evident from Fig. S6 in the Supplement. In the present of O_2, the attack of an O_2 molecule on the C-centered site of $P_{S4\text{-}add1}$, $P_{S4\text{-}add2}$, and $P_{S4\text{-}add3}$ proceed via the barrierless processes to produce the second generation peroxy radicals $P_{S4\text{-}add1}$-a/-s, $P_{S4\text{-}add2}$-a/-s and $P_{S4\text{-}add3}$-a/-s. The O_2-addition reaction occurring at the same direction as the −OOH group is defined as *syn*-O_2-addition, while the O_2-addition reaction occurring at the opposite direction as the −OOH group is defined as *anti*-O_2-addition. For the reaction $P_{S4\text{-}add1} + O_2 \rightarrow P_{S4\text{-}add1}$-a/-s, ΔE_r of *anti*-O_2-addition is −5.8 kcal mol^{-1}, which is lower than that of *syn*-O_2-addition by 0.4 kcal mol^{-1}, suggesting that *anti*-O_2-addition is preferable over *syn*-O_2-addition in energy. For the reactions $P_{S4\text{-}add2} + O_2 \rightarrow P_{S4\text{-}add2}$-a/-s and $P_{S4\text{-}add3} + O_2 \rightarrow P_{S4\text{-}add3}$-a/-s, it can be concluded the same by the ΔE_r values that *anti*-O_2-addition reaction is energetically feasible.

The resulting $P_{S4\text{-}add1}$-a/-s can proceed intramolecular cyclization reaction, where the attack of end-site oxygen atom of the −OO group on C2-site of the C_2=C_3 double bond, leading to the formation of peroxide bicyclic alkyl radicals. ΔE_a and ΔE_r of the reaction $P_{S4\text{-}add1}$-a → $P_{S4\text{-}add1}$-a-1 are 11.8 and −16.8 kcal mol^{-1}, respectively, which are lower

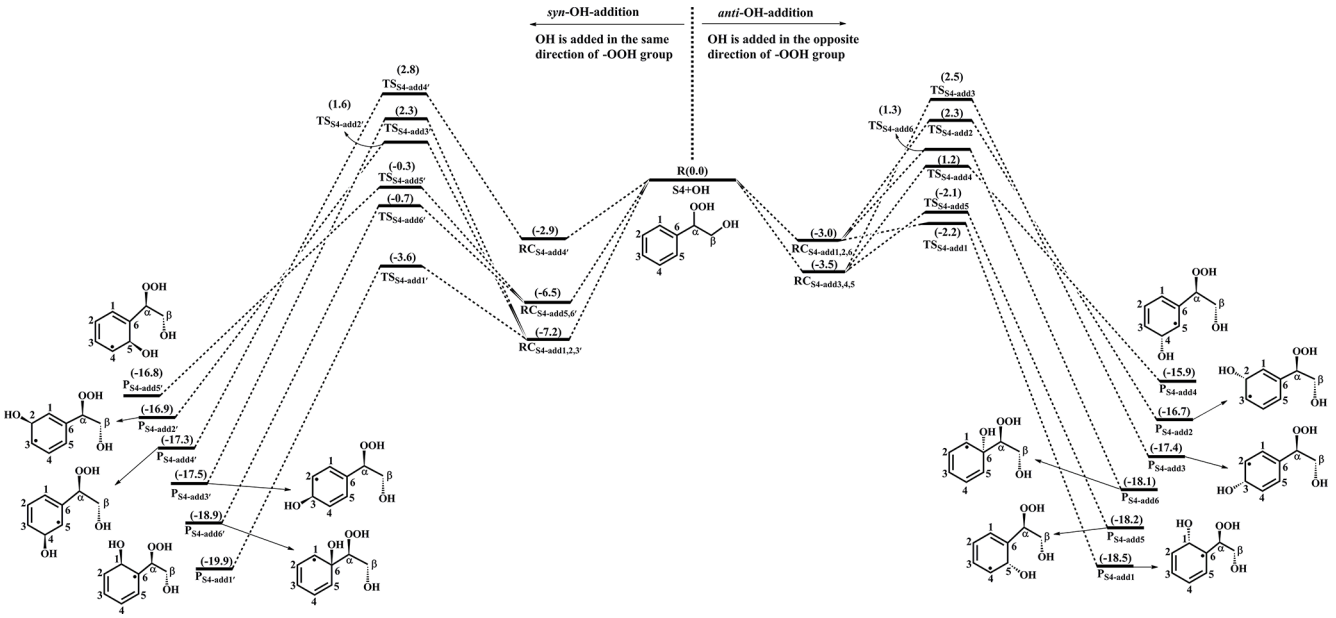

Figure 2. PES for the oxidation of 1st-ROOH(S4) initiated by OH radicals at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

than those of the reaction $P_{S4\text{-}add1}$-s → $P_{S4\text{-}add1}$-s-1 by 3.9 and 2.2 kcal mol^{-1}, respectively. The aforementioned results reveal that the intramolecular cyclization reaction of *anti*-O_2-addition product $P_{S4\text{-}add1}$-a is favorable on both thermochemically and kinetically. A similar conclusion is also derived from the intramolecular cyclization reactions of *anti*-O_2-addition products $P_{S4\text{-}add2}$-a and $P_{S4\text{-}add3}$-a. Notably, the barriers of the intramolecular cyclization reactions $P_{S4\text{-}add2}$-a → $P_{S4\text{-}add2}$-a-1 ($\Delta E_a = 31.1$ kcal mol^{-1}) and $P_{S4\text{-}add2}$-a → $P_{S4\text{-}add2}$-a-2 ($\Delta E_a = 34.6$ kcal mol^{-1}) are extremely high, making them insignificant in the atmosphere. The tautomerization between $P_{S4\text{-}add1}$-a-1 and $P_{S4\text{-}add3}$-a-1 readily occurs due to the existence of resonance structures, and it is therefore that the latter conformer is selected as a prototype for the investigating of its subsequent reaction mechanism.

The formed $P_{S4\text{-}add3}$-a-1 can combine with an O_2 molecule leading to the third generation peroxy radicals (also called as peroxide bicyclic peroxy radicals, BPR) $P_{S4\text{-}add3}$-a-2, and the lowest energy conformer is presented in Fig. S7 in the Supplement. The isomerization of $P_{S4\text{-}add3}$-a-2 may undergo through a concerted process of the cleavage of −O−O− bridge bond and C_1−C_2 bond as well as hydrogen atom transfer from the hydroxyl group to the bridge oxygen atom, yielding a new peroxy radical ($\Delta E_a = 28.5$ kcal mol^{-1}). The room temperature rate coefficient is calculated to be 3.0×10^{-9} s^{-1}, which is several orders of magnitude low than the typical pseudo-first-order rate constants $k'_{RO_2+HO_2}$ (0.01–0.02 s^{-1}) and k'_{RO_2+NO} (0.1–10 s^{-1}), suggesting that the isomerization reaction is less importance in the atmosphere. Therefore, the bimolecular reactions of $P_{S4\text{-}add3}$-a-2 with HO_2 radicals with NO are mainly taken into consideration in this study.

In the pristine environments, $P_{S4\text{-}add3}$-a-2 can react with HO_2 radicals resulting in the formation of the second generation products, bicyclic hydroperoxide 2nd-ROOH (S6) and peroxide bicyclic alkoxy radical (BAR) $P_{S4\text{-}add3}$-a-3, as depicted in Fig. S7. For the subsequent reactions of S6 initiated by OH radicals, the detailed mechanisms are discussed in Sect. 3.3.1. From Fig. 3, it can be seen that the unimolecular decomposition of $P_{S4\text{-}add3}$-a-3 involves two kinds of pathways. One is the ring-opening reaction, where the breakage of C_5−C_6 bond produces an alkyl radical S7 ($\Delta E_a = 5.9$ kcal mol^{-1}). The other is cyclization reaction, where the attack of oxygen atom of O-centered site on the C4-site of the C_3=C_4 double bond generates the ring-retaining alkyl radical S15 ($E_a = 8.0$ kcal mol^{-1}). The branching ratios for the formation of S7 and S15 are predicted to be 74.7 % and 25.3 %, respectively.

As shown in Fig. 3, S7 decomposes through the barrierless rupture of −O−O− bridge bond to form alkoxy radical S8-x, which includes five possible conformers as presented in Fig. S8 in the Supplement. The Boltzmann populations of different conformers are listed in Table S4 in the Supplement. S8-x can undergo various intramolecular H-shifts, in which a hydrogen atom is transferred from different carbon atoms to O-centered site, forming the alkyl radicals. Among the competing H-shift reactions, 1,5 H-shift occurring at the $-C_5$(O)H group exhibits the smallest barrier ($\Delta E_a = 0.6$ kcal mol^{-1}), and $k_{MC\text{-}TST}$ is calculated to be 8.2×10^9 s^{-1} at ambient temperature (Table S5 in the Supplement). The formed S8-c-P can readily isomerize to S9 due to its resonance stabilized structure. The unimolecular decom-

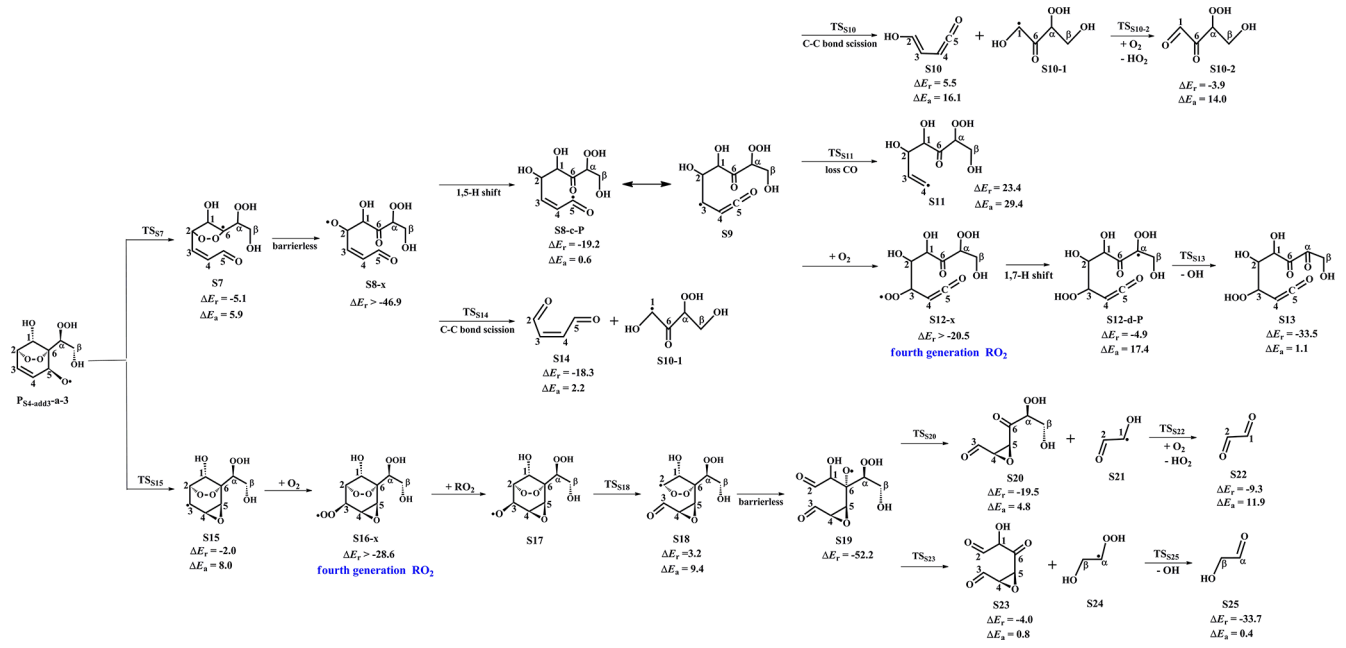

Figure 3. PES for the unimolecular decomposition of $P_{S4\text{-}add3}$-a-3 and its subsequent reactions at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

position of S9 can proceed through the C_1-C_2 bond scission to produce a ketene-enol S10 and an alkyl radical S10-1 ($\Delta E_a = 16.1\,\text{kcal mol}^{-1}$), followed by reaction with O_2 leading to a HO_2 radical and a 1,2-dicarbonyl compound S10-2 ($\Delta E_a = 14.0\,\text{kcal mol}^{-1}$). Alternatively, S9 may undergo via the elimination of CO to generate an alkyl radicals S11 ($\Delta E_a = 29.4\,\text{kcal mol}^{-1}$). The aforementioned results show that the formation of S10 and S10-1 is energetically favorable, with the rate coefficient k_{S10} of $26.1\,\text{s}^{-1}$.

In the presence of O_2, the attack of an O_2 molecule on the C-centered sites of S9 leads to the fourth generation peroxyl radical S12-x ($\Delta E_r > -20.5\,\text{kcal mol}^{-1}$). Adopting the rate coefficient k_{R+O_2} of $6.0 \times 10^{-12}\,\text{cm}^3\,\text{molec.}^{-1}\,\text{s}^{-1}$ for the reactions of alkyl radicals with O_2, and the atmospheric O_2 concentration of $5 \times 10^{18}\,\text{molec.cm}^{-3}$ (Ma et al., 2021), the pseudo-first-order rate constant $k'_{R+O_2} = k_{R+O_2}[O_2]$ is $3.0 \times 10^7\,\text{s}^{-1}$. The unimolecular decomposition of alkyl radicals is competitive only when their decay rate exceeds $3.0 \times 10^7\,\text{s}^{-1}$. k'_{R+O_2} is about six orders of magnitude greater than k_{S10}, indicating that the unimolecular decomposition of S9 is less importance. As shown in Fig. S9 in the Supplement, S12-x can proceed various intramolecular H-shift reactions, where hydrogen atom migrates from the different carbon sites or hydroxyl groups to the terminal oxygen atom of the −OO group, resulting in the formation of QOOH radicals and alkoxyl radicals. Among these competing H-shift reactions, the 1,7-H transfer at the Cα-site leading to the formation of S12-d-P exhibits the smallest barrier ($\Delta E_a = 17.4\,\text{kcal mol}^{-1}$). Then, it decomposes to yield an OH radical and a closed-shell product S13 containing a hydroperoxide, three hydroxyl and three carbonyl groups ($\Delta E_a = 1.1\,\text{kcal mol}^{-1}$).

S8-x can proceed through the C_1-C_2 bond scission to yield an unsaturated 1,4-dicarbonyl species S14 and an alkyl radical S10-1 ($\Delta E_a = 2.2\,\text{kcal mol}^{-1}$), with the rate coefficient of $2.1 \times 10^{10}\,\text{s}^{-1}$. Notably, both the 1,5 aldehyde H-shift and C_1-C_2 bond scission reactions yield a closed-shell species S10-2 with up to five oxygen atoms, and the branching ratios are predicted to be 28.1 % and 71.9 %, respectively. The result is further supported by the previous study that the proportion of aldehyde H-shift products constitutes about one third of the total products in the reaction benzene +•OH (Wang et al., 2020b).

As shown in Fig. 3, S15 can further react with O_2 leading to the fourth generation peroxy radical S16-x, which can proceed either intramolecular H-shifts forming QOOH radicals (Fig. S10 in the Supplement), or reactions with RO_2 radicals and NO forming alkoxyl radical S17. Notably, the barriers of intramolecular H-shifts are extremely high ($\Delta E_a > 34.6\,\text{kcal mol}^{-1}$), making them less importance in the atmosphere. The transformation of S17 undergoes through the breakage of C_2-C_3 bond to produce an alkyl radical S18 ($\Delta E_a = 9.4\,\text{kcal mol}^{-1}$), followed by fragmentation into an alkoxyl radical S19 via the barrierless rupture of the −O−O− bridge bond. Then, S19 dissociates to an OH radical, a glycolaldehyde S25 and a C_6-epoxide product S23 bearing a hydroxy and three carbonyl groups, being the dominant pathway. The regeneration of OH radicals drives the successive autoxidation of styrene, eventually leading to the production of multifunctional products.

3.2.2 The oxidation mechanism of 1st-$RONO_2$ initiated by OH radicals

The OH-initiated oxidation of 1st-$RONO_2$ (S5) proceeds through the addition of OH radicals to different carbon sites in the benzene ring to form various alkyl radicals $P_{S5\text{-}addx}$, as depicted in Fig. 4. Among the competing OH-addition reactions, the OH-addition reaction at the C1-site, which proceeds on the opposite direction as the $-ONO_2$ group, has the smallest barrier ($R_{S5\text{-}add1}$, $\Delta E_a = 0.4\,kcal\,mol^{-1}$) due to the stability of the formed product $P_{S5\text{-}add1}$. The result again shows that the *ortho*-addition reaction is energetically feasible. $P_{S5\text{-}add1}$ may isomerize to two other resonance structures, namely, $P_{S5\text{-}add2}$ and $P_{S5\text{-}add3}$. For the reaction $P_{S5\text{-}add1} + O_2$, O_2 may add on either the opposite (*anti*-O_2-addition) or the same direction (*syn*-O_2-addition) relative to the $-NO_3$ group, leading to the second generation peroxyl radicals $P_{S5\text{-}add1}$-a and $P_{S5\text{-}add1}$-s (Fig. S11 in the Supplement). The exoergicity of these two reactions are −6.7 and $-4.4\,kcal\,mol^{-1}$, respectively, suggesting that the *anti*-O_2-addition reaction is thermochemically favorable. Next, they can isomerize via a cyclization process to yield $P_{S5\text{-}add1}$-a-1 and $P_{S5\text{-}add1}$-s-1 with the ΔE_a of 13.3 and $18.1\,kcal\,mol^{-1}$. This result shows that the cyclization reaction of *anti*-O_2-addition product $P_{S5\text{-}add1}$-a is kinetically feasible. A similar conclusion is also obtained from the reaction $P_{S5\text{-}add3} + O_2$ that the formation of *anti*-O_2-addition product $P_{S5\text{-}add3}$-a-1 is dominant. Due to the existence of the conjugate double bond, it facilitates the tautomerization between $P_{S5\text{-}add1}$-a-1 and $P_{S5\text{-}add3}$-a-1. Therefore, we mainly focus on the subsequent chemistry of $P_{S5\text{-}add3}$-a-1 in the present study.

$P_{S5\text{-}add3}$-a-1 can further react with an O_2 molecule leading to the third generation peroxyl radicals $P_{S5\text{-}add3}$-a-2, which include multiple conformers. The lowest energy conformer resulting from conformer search is presented in Fig. S12 in the Supplement. In urban environments, the bimolecular reaction of $P_{S5\text{-}add3}$-a-2 with NO yields the second generation products, a bicyclic organic nitrate 2nd-$RONO_2$ (S26) and a BAR $P_{S5\text{-}add3}$-a-3, as displayed in Fig. S12. The detailed mechanism of OH-initiated oxidation of S26 is discussed in Sect. 3.3.2. As shown in Fig. 5, $P_{S5\text{-}add3}$-a-3 can either proceed via a ring opening process to form an alkyl radical S27 ($\Delta E_a = 7.3\,kcal\,mol^{-1}$), or undergo through a cyclization process to generate an epoxide species S35 ($\Delta E_a = 8.5\,kcal\,mol^{-1}$). The branching ratios of these two reactions are predicted to be 69.2 % and 30.8 %, respectively. Notably, the branching ratio of cyclization reaction of $P_{S5\text{-}add3}$-a-3 increases by 5.5 % compared to that of cyclization reaction of $P_{S4\text{-}add3}$-a-3, suggesting that the $-ONO_2$ substitution is beneficial to cyclization reaction.

The degradation of S27 proceeds through the barrierless scission of −O−O− bridge bond to form S28-x, and the Boltzmann populations of different conformers are listed in Table S6 in the Supplement. S28-x can undergo via various intramolecular H-shifts to produce QOOH radicals, in which hydrogen atom transfer from the −C(O)H group to the terminal oxygen atom of the −OO group forming S28-e-P has the smallest barrier ($\Delta E_a = 2.0\,kcal\,mol^{-1}$) (Fig. S13 in the Supplement). S28-e-P can readily isomerize to S29, which includes two distinct decomposition pathways. One is the C_1-C_2 bond cleavage, yielding a ketene-enol S30 and an alkyl radical S30-1 ($\Delta E_a = 17.8\,kcal\,mol^{-1}$), followed by reaction with O_2 to form a HO_2 radical and a 1,2-dicarbonyl species S30-2 ($\Delta E_a = 11.7\,kcal\,mol^{-1}$). The other is the elimination of CO to generate an alkyl radical S31 ($\Delta E_a = 24.8\,kcal\,mol^{-1}$), but the barrier is considerably high, making this pathway less competitive. The rate coefficient for the formation of S30 and S30-1 is calculated to be $14.4\,s^{-1}$, which is about six orders of magnitude lower than the pseudo-first-order rate constant k'_{R+O_2}, indicating that the unimolecular decomposition of S29 is insignificant.

In the presence of O_2, the bimolecular reaction of S29 with O_2 produces the fourth generation peroxyl radicals S32-x, comprising five energetically similar conformers as shown in Fig. S14 in the Supplement. For the 1,7-H transfer reaction, hydrogen atom at the Cα-site can be transferred through an eight-membered ring transition state to generate an alkyl radical S32-d-P ($\Delta E_a = 23.3\,kcal\,mol^{-1}$), followed by the elimination of NO_2 forming a closed product S33 ($\Delta E_a = 1.0\,kcal\,mol^{-1}$). S33 and S13 are isomeric species, with the former exhibiting more stability than the latter. S28-x can proceed through the cleavage of C_1-C_2 bond to generate an unsaturated 1,4-dicarbonyl compound S34 and an alkyl radical S30-1. The rare coefficients of the 1,5 aldehyde H-shift and C_1-C_2 bond scission reactions are predicted to be 1.7×10^9 and $5.8 \times 10^9\,s^{-1}$ (Table S7 in the Supplement), respectively, with the branching ratios of 23 % and 77 %. S30-1, formed from the above mentioned two pathways, may undergo through H-abstraction by O_2 to yield an organic nitrate S30-2 bearing a hydroxyl and two carbonyl groups ($\Delta E_a = 11.7\,kcal\,mol^{-1}$).

S35 can combine with an O_2 molecule forming the fourth generation peroxyl radicals S36-x, which have five possible conformers as shown in Fig. S15 in the Supplement. S36-x can proceed either intramolecular H-shifts forming QOOH radicals, or reaction with RO_2 radicals and NO generating alkoxyl radical S37. However, the barriers of intramolecular H-shifts are extremely high ($\Delta E_a > 31.3\,kcal\,mol^{-1}$), making them less importance in the atmosphere. The degradation of S37 initially proceeds via the breakage of C_2-C_3 bond to form S38 ($\Delta E_a = 9.2\,kcal\,mol^{-1}$), followed by decomposition into an alkoxyl radical S39 via the barrierless scission of −O−O− bridge bond. The dominant pathway of the unimolecular decomposition of S39 is the formation of a glyoxal and a C_6-epoxide species S40-1 bearing a $-NO_3$, a hydroxyl and two carbonyl groups. This process differs from the unimolecular decay of S19, where the favorable pathways is the formation of a tricarbonyl compound S23. The aforementioned results reveal that the preferable pathway is strongly dependent on the breakage of C−C bond associated with the

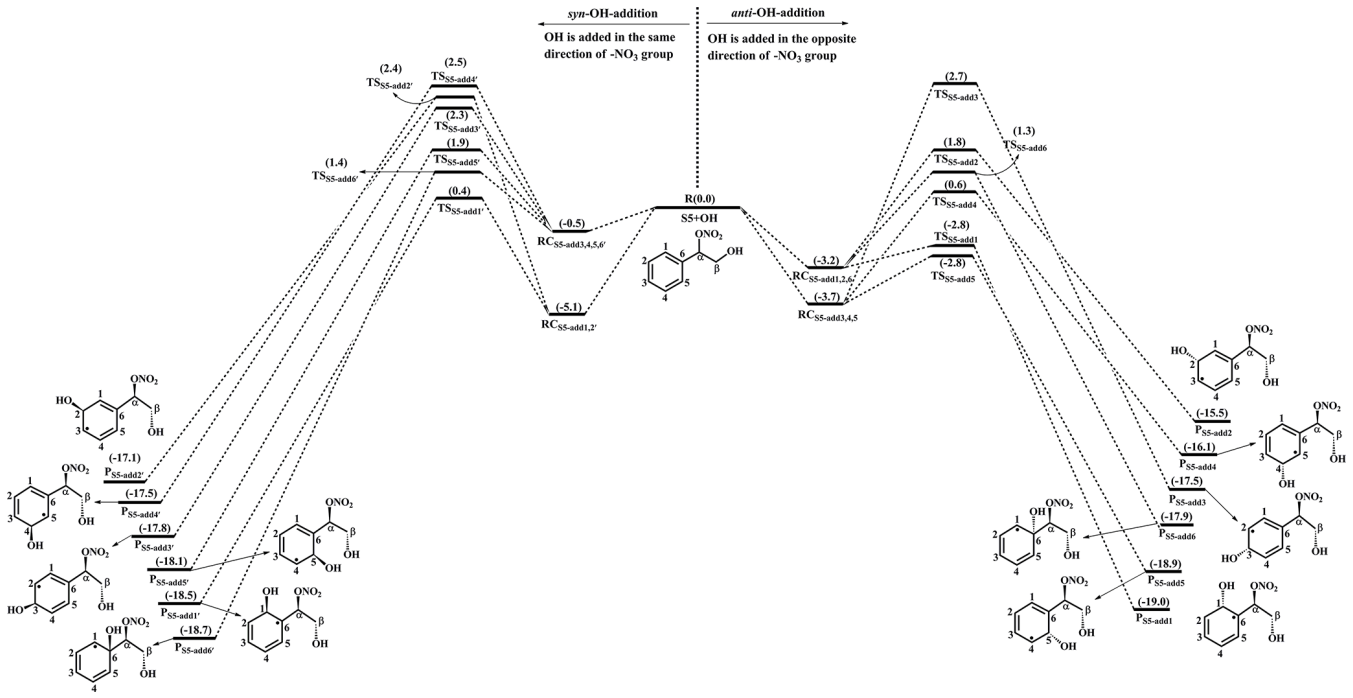

Figure 4. PES for the oxidation of 1st-RONO$_2$(S5) initiated by OH radicals at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

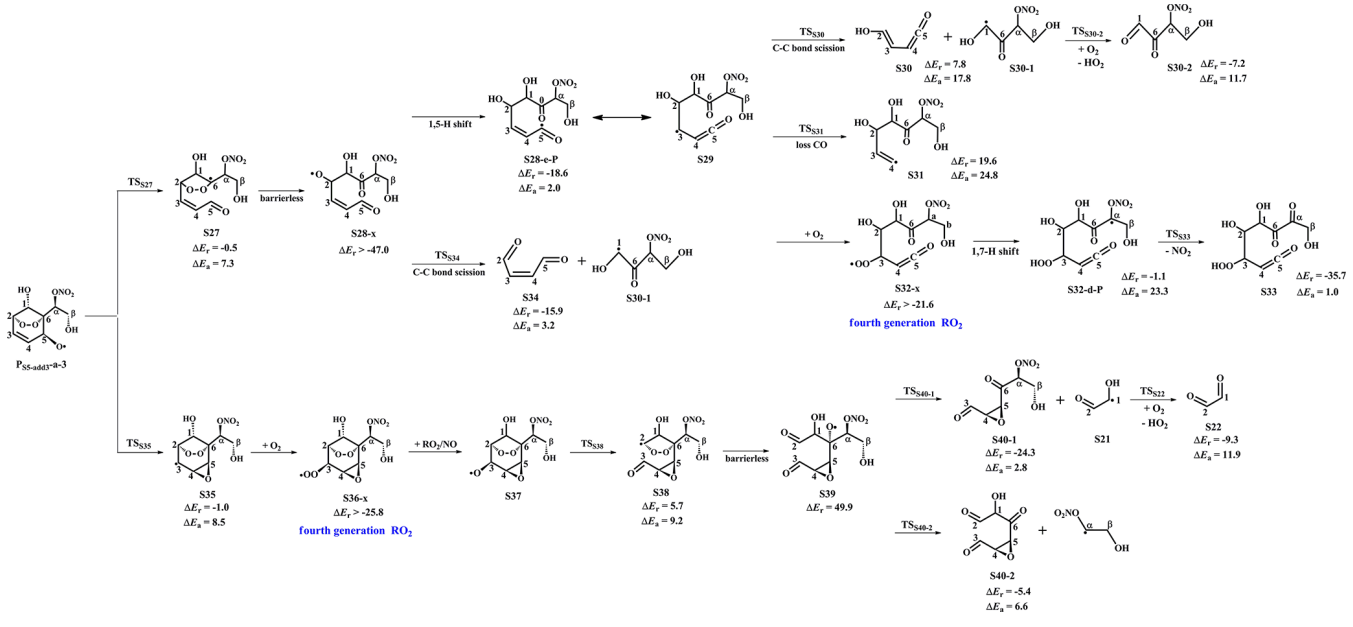

Figure 5. PES for the unimolecular decomposition of P$_{S5\text{-}add3}$-a-3 and its subsequent reactions at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

property of substituents in the decomposition of alkoxy radicals.

3.3 Third generation •OH oxidation mechanisms of 2nd-ROOH and 2nd-$RONO_2$

The second generation products, bicyclic hydroperoxide 2nd-ROOH and bicyclic organic nitrate 2nd-$RONO_2$, have multiple possible conformers. The global minimum structures of 2nd-ROOH (S6) and 2nd-$RONO_2$ (S26) resulting from the conformer search are presented in Figs. S7 and S12, respectively.

3.3.1 The oxidation mechanism of 2nd-ROOH initiated by OH radicals

OH-initiated oxidation of 2nd-ROOH (S6) can either undergo through the addition of OH radicals to either side of the C_3=C_4 double bond to generate the alkyl radicals, or proceed via H-abstraction from the different carbon sites to produce the alkyl radicals and alkoxyl radicals, as shown in Figs. S16 and S17 in the Supplement. For the OH-addition reactions, *syn*-OH-addition is defined as the addition of OH radicals on the same side as the −OOH group, while *anti*-OH-addition is referred to the addition of OH radicals on the opposite side as the −OOH group. The addition of OH radicals to the C3-site of the C_3=C_4 double bond forming the product $P_{S6\text{-}abs3}$ has the smallest barrier ($\Delta E_a = 2.4\,kcal\,mol^{-1}$) and the exoergicity of $-33.5\,kcal\,mol^{-1}$. For the H-abstraction reactions, the abstraction of hydrogen atom at the C5-site is the most favorable pathway ($\Delta E_a = 3.6\,kcal\,mol^{-1}$) and the exoergicity of $-20.2\,kcal\,mol^{-1}$. It is mainly because that the presence of an allyl group enhances the stability of the resulting product $P_{S6\text{-}abs5}$. Notably, the abstraction of hydrogen atom at the C2-site proceeds through a concerted process of C_2−H bond and −O−O− bridge bond rupture, leading to the formation of an alkoxyl radical $P_{S6\text{-}abs2}$ ($\Delta E_a = 7.2\,kcal\,mol^{-1}$). This reaction is expected to be less importance due to its higher energy barrier. The rate coefficient of the favorable OH-addition reaction is calculated to be $6.4\times10^{-11}\,cm^3\,molec.^{-1}\,s^{-1}$, which is about one order of magnitude greater than that of the preferable H-abstraction reaction ($4.1\times10^{-12}\,cm^3\,molec.^{-1}\,s^{-1}$). Based on the above discussion, it can be concluded that OH-addition reaction is favorable on both thermochemically and kinetically. This conclusion is further supported by the •OH + alkene reaction systems that OH-addition pathways are predominant (Chen et al., 2021; Yang et al., 2017; Arathala and Musah, 2024).

As depicted in Fig. S18 in the Supplement, the unimolecular decay of the product $P_{S6\text{-}add3}$ resulting from the favorable OH-addition reaction proceeds through a cyclization process to yield an epoxide compound S41 and an OH radical byproduct with the ΔE_a of $15.3\,kcal\,mol^{-1}$ and the rate coefficient k_{R41} of $1.8\times10^2\,s^{-1}$, or undergoes via intramolecular 1,4 H-shift to form a peroxy radical S42 with the ΔE_a of $21.8\,kcal\,mol^{-1}$ and the rate coefficient k_{R43} of $1.9\,s^{-1}$, or proceeds via the elimination of hydrogen atom to produce an alkene S43 with the ΔE_a of $37.9\,kcal\,mol^{-1}$. Based on the values of ΔE_a and the corresponding rate coefficients, the dominant pathway of the unimolecular decomposition of $P_{S6\text{-}add3}$ is the formation of S41. In the presence of O_2, the pseudo-first-order rate constant k'_{R+O_2} of the reactions of alkyl radicals with O_2 is $3.0\times10^7\,s^{-1}$, which is about five orders of magnitude greater than k_{R41}, suggesting that the unimolecular decomposition of $P_{S6\text{-}add3}$ is insignificant.

As shown in Fig. 6, the fourth generation peroxy radicals S44-x formed in the addition reaction $P_{S6\text{-}add3} + O_2$ can either proceed via intramolecular H-shits to form QOOH, or undergo through self- or cross-reactions to yield an alkoxy radical S45. Due to the considerably high barriers of intramolecular H-shifts, they are deemed to be negligible under atmospheric conditions. S45 can convert into an alkyl radical S46 through the cleavage of C_4−C_5 bond, or dissociate to an alkyl radical S50 via the rupture of C_3−C_4 bond. The barrier of the former reaction is $3.9\,kcal\,mol^{-1}$, which is lower than that of the latter pathway by $2.6\,kcal\,mol^{-1}$, indicating that the formation of S46 is kinetically preferable. Then, S46 decomposes into an OH radical byproduct and a C_8-product S47 bearing a −OOH, a peroxide bridge, two carbonyls, and three hydroxy groups, which is expected to be the dominant pathway owing to its lower barrier. The rate coefficient k_{RS47} is estimated to be $1.8\times10^9\,s^{-1}$, which is about two orders of magnitude greater than the pseudo-first-order rate constant k'_{R+O_2} ($3.0\times10^7\,s^{-1}$). The result reveals that the unimolecular decomposition of S46 is more competitive than the bimolecular reaction with O_2. The formed OH radicals can once again participate in the oxidations of styrene and its multifunctional products, continuing these processes until they are completely consumed.

3.3.2 The oxidation mechanism of 2nd-$RONO_2$ initiated by OH radicals

OH-initiated oxidation of 2nd-$RONO_2$ (S26) includes four different OH-addition pathways and five different H-abstraction pathways, as displayed in Figs. S19 and S20 in the Supplement. For the OH-addition reactions, the attack of OH radicals on the C3-site of the C_3=C_4 double bond forming the product $P_{S26\text{-}add3}$, occurring on the same direction relative to the $-ONO_2$ group, is found to be the favorable pathway ($\Delta E_a = 2.4\,kcal\,mol^{-1}$, $\Delta E_r = -33.6\,kcal\,mol^{-1}$). For the H-abstraction reactions, the abstraction of hydrogen atom at the C5-site is identified as the preferable pathway ($\Delta E_a = 5.7\,kcal\,mol^{-1}$, $\Delta E_r = -20.1\,kcal\,mol^{-1}$) due to the enhanced stability of the resulting product $P_{S26\text{-}add5}$ by the presence of an allyl group. By comparing the values of ΔE_a and ΔE_r of the favorable OH-addition and H-abstraction pathways, it can concluded that the former case is dominant on both thermochemically and kinetically. This conclusion is consistent with the result from the reaction

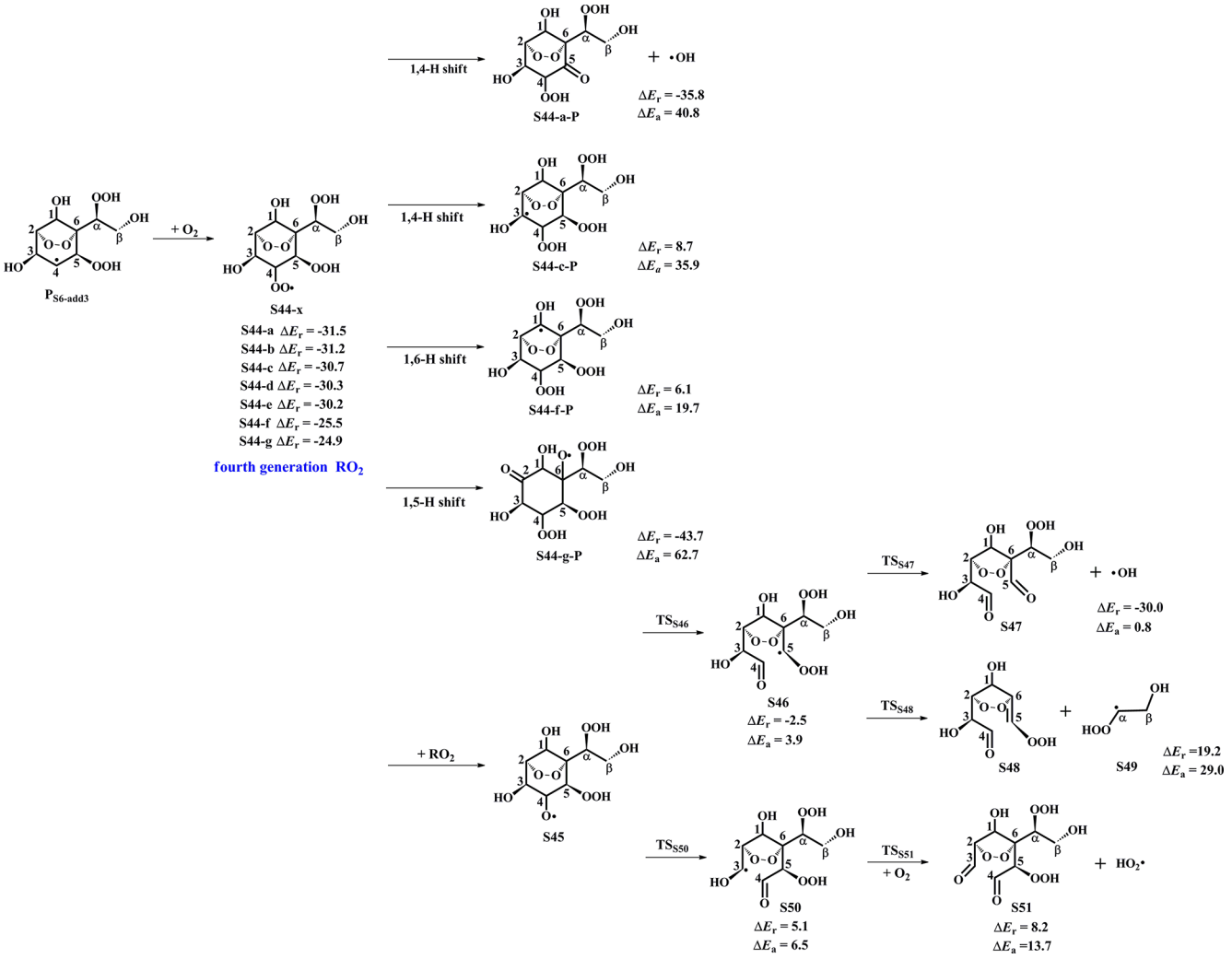

Figure 6. PES for the subsequent reactions of $P_{S6\text{-}add3}$ in the presence of O_2 at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

2nd-ROOH (S6) +•OH that OH-addition is more competitive than H-abstraction.

The product $P_{S26\text{-}add3}$ arising from the favorable OH-addition pathway has three potential unimolecular decay pathways, as depicted in Fig. S21 in the Supplement: (1) $P_{S26\text{-}add3}$ dissociates to an epoxide S52 and a NO_2 molecule through a cyclization process with the ΔE_a of 18.5 kcal mol^{-1} and the rate coefficient k_{R52} of 0.4 s^{-1}; (2) $P_{S26\text{-}add3}$ isomerizes to an alkyl radical S53 via the intramolecular 1,2 H-shift ($\Delta E_a = 40.0$ kcal mol^{-1}); (3) $P_{S26\text{-}add3}$ converts into an alkene S54 via the elimination of hydrogen atom ($\Delta E_a = 39.1$ kcal mol^{-1}). Based on the value of ΔE_a and the corresponding rate coefficient, the dominant pathway of the unimolecular decomposition of $P_{S26\text{-}add3}$ is the formation of S52. k_{R52} is about seven orders of magnitude lower than the pseudo-first-order rate constant k'_{R+O_2}, indicating that the unimolecular decomposition of $P_{S26\text{-}add3}$ is less importance.

In the presence of O_2, $P_{S26\text{-}add3}$ can react with an O_2 molecule leading to the formation of the fourth generation peroxy radicals S55-x, comprising seven possible conformers as shown in Fig. 7. For the intramolecular H-shifts of S55-x, not all of reactants (S55-c, S55-d and S55-e) have the suitable conformers that allow for the pathways across the reaction barriers. The barriers of intramolecular H-shifts are considerably high ($\Delta E_a = 20.8$ kcal mol^{-1}), making them uncompetitive in the atmosphere. Alternatively, S55-x can react with other RO_2 radicals forming an alkoxyl radical S56, followed by decomposition into an alkyl radical S57 via the breakage of C_4-C_5 bond ($\Delta E_a = 2.0$ kcal mol^{-1}), or fragmentation into an alkyl radical S61 through the cleavage of C_3-C_4 bond ($\Delta E_a = 4.5$ kcal mol^{-1}). The aforementioned results reveal that the formation of S57 is energetically favorable, which is consistent with the conclusion derived from the unimolecular decomposition of S45 that the breakage of C_4-C_5 bond is feasible. Next, S57 dissociates

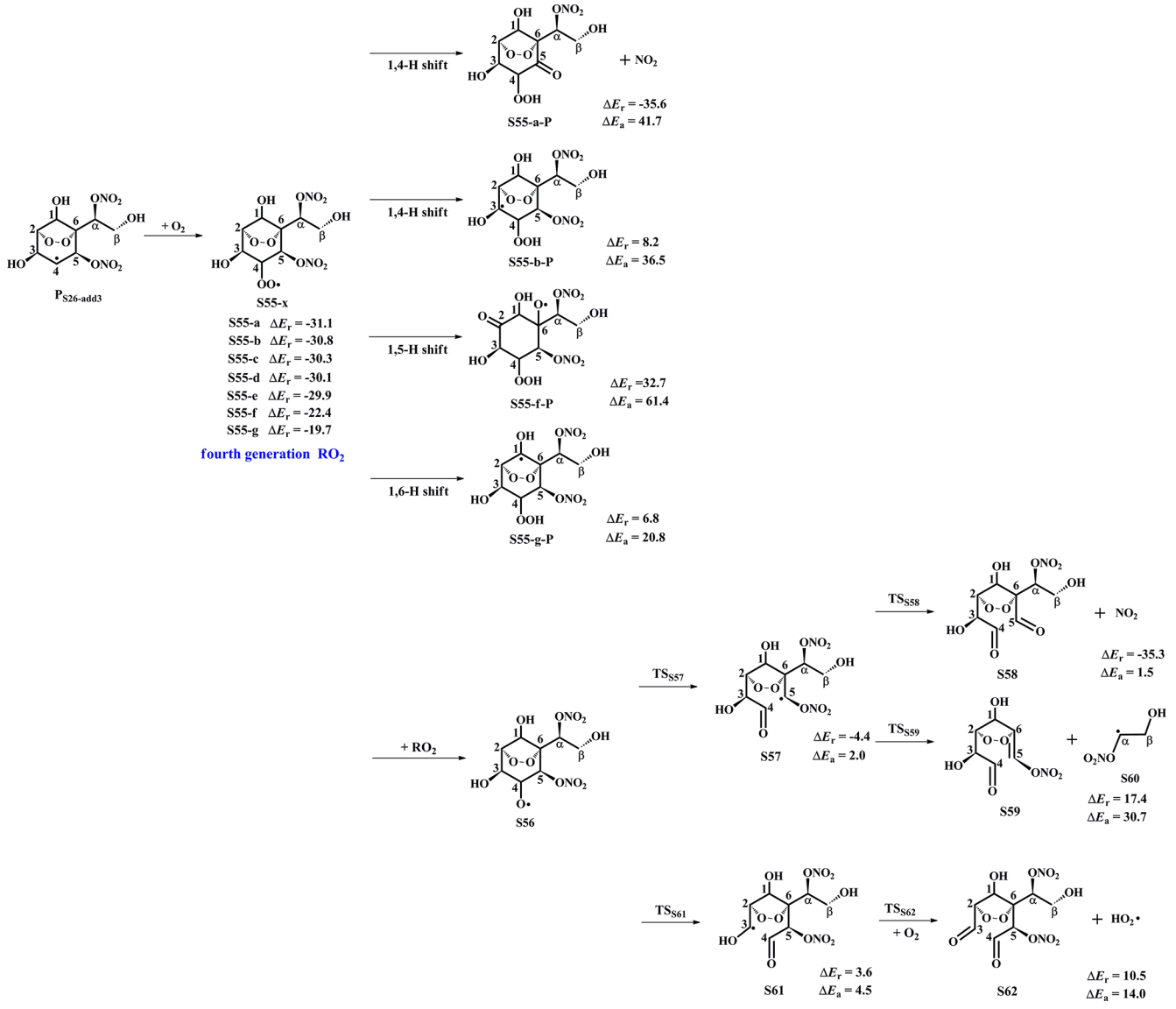

Figure 7. PES for the subsequent reactions of $P_{S26\text{-}add3}$ in the presence of O_2 at the M06-2X/6-311++G(3df,3pd)//M06-2X/6-31+g(d,p) level.

to a NO_2 coproduct and a C_8-product S58 that possessed a $-NO_3$, a peroxide bridge, two carbonyls, and three hydroxy groups. This pathway is expected to be the dominant one ($\Delta E_a = 1.5\,\text{kcal mol}^{-1}$), with the rate coefficient k_{RS58} of $1.2 \times 10^9\ \text{s}^{-1}$. The resulting NO_2 can further participate in the cycling of NO_x, ultimately generating tropospheric ozone and SOA.

The overall reaction mechanism and the fractional yields of the major products in the multi-generation •OH oxidation of styrene under different NO_x conditions are presented in Figs. S22 and S23 in the Supplement. In the low-NO_x conditions, the fractional yield of the first generation closed-shell product 1st-ROOH (S4) formed from the reaction S2-1-x + HO_2• is predicted to be 71.6 %. For the second generation •OH oxidation, the reaction of the peroxyl radical $P_{S4\text{-}add3}$-a-2 with HO_2 radicals produces the second generation closed-shell product 2nd-ROOH (S6) and an alkoxyl radical $P_{S4\text{-}add3}$-a-3, with the fractional yields of 41.4 % and 10.4 %, respectively. The formed $P_{S4\text{-}add3}$-a-2 can either proceed through the C_5-C_6 bond scission to produce an alkyl radical S7 with the fractional yield of 7.8 %, or undergo via a cyclization process to generate an alkyl radical S15 with the fractional yield of 2.6 %. S7 and S15 can be transformed via a series of reactions, ultimately leading to the formation of second generation closed-shell product S10-2, S13 and S23, with the fractional yields of 5.6 %, 2.2 % and 1.3 %, respectively. For the third generation •OH oxidation, the degradation of 2nd-ROOH (S6) ultimately yields the third genera-

tion closed-shell products S47 and S51, with the fractional yields of 26.3 % and 0.3 %, respectively. As a result, the major closed-shell products are 1st-ROOH (S4), 2nd-ROOH (S6), S10-2, S13 and S47 in the multi-generation •OH oxidation of styrene in the low-NO_x conditions.

In the high-NO_x conditions, the fractional yield of the first generation closed-shell product 1st-$RONO_2$ (S5) formed from the reaction S2-1-x + NO is predicted to be 26.5 %, as shown in Fig. S23. As the •OH oxidation reactions proceed, 1st-$RONO_2$ (S5) can be initially transformed into the peroxyl radical $P_{S5\text{-}add3}$-a-2, followed by reaction with NO to form the second generation closed-shell product 2nd-$RONO_2$ (S26) and an alkoxyl radical $P_{S5\text{-}add3}$-a-3, with the fractional yields of 4.8 % and 11.2 %, respectively. The decomposition of $P_{S5\text{-}add3}$-a-3 undergoes via two distinct pathways. One is the C_5-C_6 bond cleavage, leading to an alkyl radical S27 with the fractional yield of 7.8 %. The other is the cyclization, resulting in an alkyl radical S35 with the fractional yield of 3.4 %. The resulting S27 and S35 undergo multiple oxidation steps, finally leading to the formation of the second generation closed-shell products S30-2, S33 and S40-1, with the fractional yields of 6.0 %, 1.8 %, and 1.7 %, respectively. 2nd-$RONO_2$ (S26) can be further oxidized to yield the third generation closed-shell products S58 and S62, with the fractional yields of 2.6 % and 0.03 %, respectively. In summary, the major closed-shell products are 1st-$RONO_2$ (S5), 2nd-$RONO_2$ (S26), S30-2 and S58 in the multi-generation •OH oxidation of styrene in the high-NO_x conditions.

3.4 Volatility classes

The volatility classes for various organic compounds are based on their saturation concentration, as proposed by Donahue et al. (2012). The saturated vapour pressure (P^0) and saturated concentration (c^0) of styrene and its multi-generation •OH oxidation products are predicted by using the SIMPOL.1 method (Pankow et al., 2008). As show in Table S8 in the Supplement, the P^0 and c^0 of the first generation closed-shell product benzaldehyde (C_7H_6O) are 7.62×10^{-4} atm and $2.89 \times 10^6\ \mathrm{ug\,m^{-3}}$, respectively, which are 3–4 orders of magnitude greater than those of S4 ($C_8H_{10}O_3$, $P^0 = 1.43 \times 10^{-7}$ atm and $c^0 = 8.89 \times 10^2\ \mathrm{ug\,m^{-3}}$) and S5 ($C_8H_9NO_3$, $P^0 = 2.54 \times 10^{-7}$ atm and $c^0 = 1.87 \times 10^3\ \mathrm{ug\,m^{-3}}$). Based on the values of c^0, benzaldehyde is classified as the volatile organic compounds (VOCs), whereas S4 and S5 are classified as the intermediate volatility organic compounds (IVOCs). These first generation closed-shell products exist exclusively in the gas phase under atmospheric conditions (Bianchi et al., 2019).

For the second generation closed-shell products, S6 ($C_8H_{12}O_8$, $c^0 = 4.50 \times 10^{-2}\ \mathrm{ug\,m^{-3}}$) and S26 ($C_8H_{10}N_2O_{10}$, $c^0 = 0.18\ \mathrm{ug\,m^{-3}}$) formed from the bimolecular reactions with HO_2 radicals and NO are classified as the low volatility organic compounds (LVOCs). Similarly, S13 ($C_8H_{10}O_8$, $c^0 = 2.97 \times 10^{-2}\ \mathrm{ug\,m^{-3}}$) and S33 ($C_8H_{10}O_8$, $c^0 = 2.97 \times 10^{-2}\ \mathrm{ug\,m^{-3}}$), formed through the ring-opening and subsequent intramolecular H-shift reactions of $P_{S4\text{-}add3}$-a-3 and $P_{S5\text{-}add3}$-a-3, respectively, are also classified as LVOCs, which can condense onto the existing large particles (Bianchi et al., 2019). The c^0 values of the remaining closed-shell products are significantly greater than those of the aforementioned four products, for example, the c^0 values of S20 ($C_6H_8O_6$) and S40-1 ($C_6H_7NO_7$), formed by the cyclization and decomposition reactions of $P_{S4\text{-}add3}$-a-3 and $P_{S5\text{-}add3}$-a-3, are 42.21 and 75.86 $\mathrm{ug\,m^{-3}}$, respectively, classifying them as the semivolatile organic compounds (SVOC).

For the third generation closed-shell products, the c^0 values of S47 ($C_8H_{12}O_9$, $c^0 = 2.68 \times 10^{-4}\ \mathrm{ug\,m^{-3}}$) and S51 ($C_8H_{10}O_{10}$, $c^0 = 1.58 \times 10^{-4}\ \mathrm{ug\,m^{-3}}$), formed through the O_2-addition and subsequent decomposition reactions of $P_{S6\text{-}add3}$, are about two orders of magnitude lower than those of the second generation closed-shell products S6 and S13, despite being classified as LVOCs. Similarly, S58 ($C_8H_{11}NO_{10}$, $c^0 = 5.37 \times 10^{-4}\ \mathrm{ug\,m^{-3}}$) and S62 ($C_8H_{10}N_2O_{12}$, $c^0 = 6.18 \times 10^{-4}\ \mathrm{ug\,m^{-3}}$), formed via the O_2-addition and subsequent decomposition reactions of $P_{S26\text{-}add3}$, exhibit lower c^0 values compared to the second generation closed-shell products S26 and S33. The aforementioned results reveal that the volatility of the multi-generation •OH oxidation products significantly decreases with increasing the number of •OH oxidation steps. As the oxidation reactions of the third generation closed-shell products proceed further, the formed products may possess sufficiently low volatility to participate in the formation and growth of new aerosol particle.

4 Conclusions and atmospheric implications

The results reveal that the first generation RO_2 radicals, formed from the addition of OH radicals to the C_β-site of a vinyl group in styrene followed by O_2-addition, can proceed intramolecular H-shifts to generate various alkyl and alkoxyl radicals. The rate coefficient $k_{\text{MC-TST}}$ is calculated to be $1.6 \times 10^{-4}\ \mathrm{s^{-1}}$. Among the competing H-shift pathways, the hydrogen atom transfer from the −OH group to the terminal oxygen atom of the −OO group has the lowest barrier. The resulting alkoxy radical can further decompose into benzaldehyde through the successive elimination of HCHO and an OH radical. The 1,5-H shift reaction occurring at the −OH group is the rate-determining step in the formation of benzaldehyde. Alternatively, the first generation RO_2 radicals can proceed bimolecular reactions with HO_2 radicals and NO, leading to the formation of the first generation closed-shell C7- and C8-products 1st-ROOH ($C_8H_{10}O_3$), benzaldehyde (C_7H_6O), and 1st-$RONO_2$ ($C_8H_9NO_3$).

For the second generation •OH oxidation, OH-addition reaction occurring at the *ortho*-site of 1st-ROOH and 1st-$RONO_2$ has a significant dominance. This is consistent with

the analogous reaction systems, toluene +•OH and phenol +•OH, in which *ortho*-OH-addition reaction is energetically favorable (Wu et al., 2020; Xu and Wang, 2013). The resulting alkyl radicals may undergo two O_2-addition steps and a cyclization process to form BPR, which can react with HO_2 radicals and NO to yield the corresponding BAR, and the second generation closed-shell C8-product 2nd-ROOH ($C_8H_{12}O_8$) and 2nd-RONO$_2$ ($C_8H_{10}N_2O_{10}$), with the fractional yields of 41.4 % and 4.8 %. The unimolecular decomposition of BAR formed in the reaction 1st-ROOH +•OH includes two distinct pathways: (1) ring-opening and followed by decomposition, yielding the multifunctional products S10-2 ($C_4H_6O_5$) and S13 ($C_8H_{10}O_8$) with the fractional yields of 5.6 % and 2.2 %, respectively; or (2) cyclization and followed by dissociation, generating the closed-shell C6-product S23 ($C_6H_6O_5$) with the fractional yield of 1.3 %. The major products formed from the unimolecular decomposition of BAR in the reaction 1st-RONO$_2$ + •OH are the multifunctional products S30-2 ($C_4H_5NO_6$), S33 ($C_8H_{10}O_8$) and S40-1 ($C_6H_7NO_7$), with the fractional yields of 6.0 %, 1.8 % and 1.7 %, respectively.

For the third generation •OH oxidation, the addition of OH radicals to the C=C bond in 2nd-ROOH and 2nd-RONO$_2$ is the dominant pathway. The resulting alkyl radicals can proceed a series of reactions to produce the alkoxyl radicals, which subsequently decompose into an OH radical byproduct and a closed-shell C_8-product S47 ($C_8H_{12}O_9$), identified as the favorable pathway in the reaction 2nd-ROOH +•OH. S47 contains a −OOH, a peroxide bridge, two carbonyls, and three hydroxy groups. The major product formed in the reaction 2nd-RONO$_2$ + •OH is a closed-shell C_8-product S58 ($C_8H_{11}NO_10$), which contain a $-NO_3$, a peroxide bridge, two carbonyls, and three hydroxy groups. The fractional yields of S47 and S58 are 26.3 % and 2.6 %, respectively. The volatility of the oxidation products significantly decreases with increasing the number of •OH oxidation steps in the multi-generation •OH oxidation of styrene.

In the laboratory chamber experiments, the structures of some specific oxidation products remain uncharacterized but are merely inferred from the exact mass and fragmentation data. Using high-level quantum chemistry methods, we identify the molecular structures of multifunctional products and elucidate their formation pathways in the multi-generation •OH oxidation of styrene. The mechanistic insights derived from this work are broadly applicable to the photooxidation of structurally analogous aromatics. Furthermore, we quantify the yields of multifunctional products and demonstrate that their volatility decreases significantly with increasing the number of •OH oxidation steps. The resulting multifunctional products may undergo a series of oxidation reactions to form low volatility compounds, thereby contributing to the formation and growth of new aerosol particle. In the future, more detailed experimental and theoretical studies need to be conducted to identify the molecular structures and formation pathways of multifunctional products formed through the photooxidation of other aromatics under both low and high-NO_x conditions. These studies will facilitate a more accurate characterization of the chemical composition and formation yields of aromatic SOA, and thereby help narrow the gap between the measured and modeled SOA concentrations in urban environments.

Data availability. Datasets are accessible by contacting the corresponding author, Yu Huang, (huangyu@ieecas.cn).

Supplement. Tables S1 and S3 list the energy barriers of all the elementary reactions involved in the addition of OH radicals to styrene and 1st-ROOH (S4) predicted at different levels. Tables S2, S4 and S6 list the relative electronic energy, free energy and Boltzmann population of different conformers involved in S2-1-x, S8-x and S28-x. Tables S5 and S7 list the MC-TST rate coefficients for the intramolecular H-shift reactions of S8-x and S28-x. Table S8 summaries the saturated vapour pressure and saturated concentrations of styrene and its multiple generation •OH oxidation closed-shell products. Figures S1–S3 display the PESs for the unimolecular reactions of S2-2-x, S2-3-x and S2-4-x. Figure S4 shows the global minimum structures of 1st-ROOH(S4) and 1st-RONO$_2$(S5). Figure S5 depicts the geometric parameters of toluene and 1st-ROOH (S4) and the NPA atomic charges of all the carbon atoms. Figures S6 and S11 show the PESs for the addition reactions $P_{S4\text{-}add1} + O_2$ and $P_{S5\text{-}add1} + O_2$. Figures S7 and S12 present the lowest energy conformers of third generation peroxy radicals $P_{S4\text{-}add3}$-a-2 and $P_{S5\text{-}add3}$-a-2. Figures S8–S10 depict the PESs for the intramolecular hydrogen transfer reactions of S8-x, S12-x and S16-x. Figures S13–S15 depict the PESs for the intramolecular hydrogen transfer reactions of S28-x, S32-x and S36-x. Figures S16–18 show the PESs for the OH-initiated oxidation of 2nd-ROOH (S6) and unimolecular decomposition of $P_{S6\text{-}add3}$. Figures S19–S21 show the PESs for the OH-initiated oxidation of 2nd-RONO$_2$ (S26) and unimolecular decomposition of $P_{S26\text{-}add3}$. Figures S22 and S23 show the overall reaction mechanism of the multi-generation •OH oxidation of styrene in the low- and high-NO_x conditions. The supplement related to this article is available online at https://doi.org/10.5194/acp-26-4823-2026-supplement.

Author contributions. LC and YH conceptualized the study. LC conducted quantum chemical calculation. YX and ZJ analyzed the data. LC conducted the volatility estimation. All authors discussed the results and commented on the manuscript.

Competing interests. The contact author has declared that none of the authors has any competing interests.

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