



# Chlorine radical-initiated atmospheric oxidation of imines: implications for structural influence on the nitrosamine formation

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Received: 3 October 2025 – Discussion started: 22 October 2025

Revised: 18 December 2025 – Accepted: 11 January 2026 – Published: 3 March 2026

**Abstract.** Chlorine radical ( $\bullet\text{Cl}$ ) initiated atmospheric oxidation of organic nitrogen compounds (ONCs) plays an important role in carcinogenic nitrosamines formation. Imines are important constituents of ONCs, primarily formed from the atmospheric oxidation of amines. However,  $\bullet\text{Cl}$ -initiated atmospheric oxidation of imines remains poorly understood. Here, we studied the reaction mechanisms and kinetics of  $\bullet\text{Cl}$ -initiated atmospheric oxidation for five representative imines ( $\text{CH}_2=\text{NH}$ ,  $\text{CH}_3\text{CH}=\text{NH}$ ,  $\text{CH}_3\text{N}=\text{CH}_2$ ,  $(\text{CH}_3)_2\text{C}=\text{NH}$ ,  $\text{HN}=\text{CHCH}_2\text{OH}$ ) to elucidate their atmospheric fate and extend the limited available data of ONCs, thereby establishing a structure-activity relationship for the reactions. The calculated overall reaction rate constants ( $\times 10^{-11} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ ) of  $\bullet\text{Cl} + \text{CH}_2=\text{NH}$ ,  $\bullet\text{Cl} + \text{CH}_3\text{CH}=\text{NH}$ ,  $\bullet\text{Cl} + \text{CH}_3\text{N}=\text{CH}_2$ ,  $\bullet\text{Cl} + (\text{CH}_3)_2\text{C}=\text{NH}$ , and  $\bullet\text{Cl} + \text{HN}=\text{CHCH}_2\text{OH}$  are 4.5, 27.2, 7.32, 44.8 and 12.6, respectively, which are consistent with the available experimental values. Importantly, our results show that  $\bullet\text{Cl}$ -initiated reactions of the NH-containing imines mainly produce N-centered radicals. These N-centered radicals exhibit various fates under tropospheric conditions: mainly reacting with NO to form nitrosamines or with  $\text{O}_2$  to form cyanide compounds, which differs substantially from the behavior of previously reported amines. The various fates of the N-centered radicals formed from imines originates from the difference in direct hydrogen abstraction reaction rate constants ( $k_{\text{O}_2}$ ) with  $\text{O}_2$  and the reaction rate ( $k_{\text{NO}}$ ) with NO, both of which are principally governed by the distinct molecular structure of N-centered radicals. The revealed reaction mechanism provides new insights into the atmospheric transformation and risks of imines, and enrich our understanding of  $\bullet\text{Cl}$  and ONCs chemistry.

## 1 Introduction

Chlorine radicals ( $\bullet\text{Cl}$ ) are key atmospheric oxidants that significantly impact air quality and radical budgets by oxidizing volatile organic compounds (VOCs), thereby contributing to the formation of  $\text{O}_3$ ,  $\bullet\text{OH}$  and secondary organic aerosol (SOA), and consequent impacts on climate (Gunthe et al., 2021; Liu et al., 2024; Wang et al., 2019; Yi et al., 2023; Cao et al., 2024; Ma et al., 2025). Historically,  $\bullet\text{Cl}$  was considered to originate mainly from heterogeneous reactions on

sea salt particles (Faxon and Allen, 2013). In recent years, diverse sources of  $\bullet\text{Cl}$  precursors have been identified in urban and suburban environments, e.g., coal combustion, biomass burning, road salt application, chlorine-based fertilizers, and automotive braking processes (Li et al., 2020; Fu et al., 2018; Li et al., 2024; Yin et al., 2022; Thornton et al., 2010; Cooke et al., 2025; Chen et al., 2025a). Daytime  $\bullet\text{Cl}$  concentrations have been reported to reach up to  $10^6 \text{ molecule cm}^{-3}$ , comparable to typical  $\bullet\text{OH}$  concentrations (Young et al., 2014; Li et al., 2025; Wang et al., 2023). Laboratory and theoretical

studies have shown that  $\bullet\text{Cl}$ -initiated oxidation of VOCs proceeds 10–100 times faster than those initiated by  $\bullet\text{OH}$  (Edwards and Young, 2024). Moreover,  $\bullet\text{Cl}$ -initiated oxidation of VOCs often follows distinct pathways and produce products that differ significantly from those of  $\bullet\text{OH}$ -initiated reactions (Chen et al., 2025b; Guo et al., 2020; Wang et al., 2022). For example,  $\bullet\text{Cl}$ -initiated oxidation of VOCs could yield chlorine-containing products, potentially altering their toxicity (Fu et al., 2024; Ding et al., 2021). Thus,  $\bullet\text{Cl}$ -initiated reactions represent a crucial role in governing the fate of VOCs and warrant careful consideration.

Organic nitrogen compounds (ONCs), a major subclass of VOCs, account for approximately 65 % of VOCs and are commonly detected in the atmosphere (Abudumutailifu et al., 2024). ONCs play significant roles in the formation of particles and hazardous substances, such as hydrogen cyanide (HCN), nitrous oxide ( $\text{N}_2\text{O}$ ), nitrosamines, and nitramines (Ning et al., 2022; Wang et al., 2023; Liu et al., 2021; Abudumutailifu et al., 2024; Sun et al., 2024; Elm et al., 2016). Imines ( $\text{R}_1\text{R}_2\text{C}=\text{NR}_3$ ) comprise about 16 % of atmospheric ONCs (Ditto et al., 2022), which originate from multiple sources, including combustion, motor vehicles, cleaning sports equipment and the atmospheric oxidation of amines, the latter being the primary one (Waterman and Hillhouse, 2008; Zhu et al., 2022; You et al., 2022; Onel et al., 2014). In urban areas, imines have been detected at ppt levels, comparable to typical concentrations of amines (Zhu et al., 2022). Given their abundance and their role as first-generation oxidation products of amines, elucidating the atmospheric reaction mechanisms of imines is essential for accurately assessing their environmental impacts.

To date, only limited studies have investigated the atmospheric chemistry of imines (Ali, 2020; Xu et al., 2024; Ali et al., 2016; Bunkan et al., 2014; Ali et al., 2018; Bunkan et al., 2022; Almeida and Kurtén, 2022; Ditto et al., 2020; Yao et al., 2016). Atmospheric oxidation by  $\bullet\text{OH}$  and  $\bullet\text{Cl}$  are important removal pathways for imines. Previous experimental and theoretical studies have primarily examined the reaction mechanism and kinetics of  $\bullet\text{OH}$ -initiated reactions with simple imines, such as methylimine ( $\text{CH}_2=\text{NH}$ ) (Bunkan et al., 2014; Ali et al., 2018), N-methylmethylimine ( $\text{CH}_3\text{N}=\text{CH}_2$ ) (Bunkan et al., 2022), and cyclic imine 1,2,3,6-tetrahydropyridazine (THPyz) (Almeida and Kurtén, 2022). These studies demonstrated that the substituents on the  $\text{C}=\text{N}$  bond profoundly influence the reactivity of imines, ultimately leading to varying atmospheric impacts. Nevertheless, no previous studies have addressed  $\bullet\text{Cl}$ -initiated reactions of imines. Since  $\bullet\text{OH}$  and  $\bullet\text{Cl}$ -initiated reactions of ONCs follow distinct mechanisms and yielding different products (Xie et al., 2015; Xue et al., 2022; Xie et al., 2014), the reactivity of imines toward  $\bullet\text{Cl}$  cannot be directly inferred from their  $\bullet\text{OH}$  chemistry.

Our previous studies demonstrated that  $\bullet\text{Cl}$  exhibits a particularly strong interaction with the  $-\text{NH}_x$  group of ONCs, leading to the formation of N-centered radicals, which are

recognized precursors of carcinogenic nitrosamines (Xie et al., 2017; Liu et al., 2019). Additionally, the yield of N-centered radicals is strongly dependent on the structure of ONCs. Given the electronic structures of imines,  $\bullet\text{Cl}$ -initiated atmospheric oxidation of imines has the potential to form N-centered radicals. However, the nitrogen lone pair electrons of imines can be delocalized to some extent to their adjacent  $\text{C}=\text{N}$  bond, which tends to decrease its electron-donor ability compared with previously well-studied ONCs (Carey and Sundberg, 2007), which may alter the reaction mechanisms of  $\bullet\text{Cl} + \text{imines}$  reactions. Moreover, the subsequent reactions of N-centered radicals depended on their specific structures. Therefore, to broaden our comprehension of the atmospheric chemistry of imines and fully assess their environmental risks, it is essential to elucidate the reaction mechanism and kinetics of  $\bullet\text{Cl} + \text{imines}$  reactions.

In this study, the reaction mechanisms and kinetics of imines with  $\bullet\text{Cl}$  were investigated by selecting five simple imines {i.e., methylimine ( $\text{CH}_2=\text{NH}$ ), ethanimine ( $\text{CH}_3\text{CH}=\text{NH}$ ), N-methylmethanimine ( $\text{CH}_3\text{N}=\text{CH}_2$ ), 2-propanimine ( $(\text{CH}_3)_2\text{C}=\text{NH}$ ), and 2-iminoethanol ( $\text{HN}=\text{CHCH}_2\text{OH}$ )} as model compounds. A combination of quantum chemical calculations and kinetic modeling was employed to elucidate these reactions. The parent amines of all selected imines have been detected in the ambient atmosphere (Qiu and Zhang, 2013; Ge et al., 2011; Shen et al., 2023). For the imine radicals yielded in the initial reaction step, their subsequent reactions including isomerization/dissociation, and bimolecular reactions with key atmospheric oxidants ( $\text{O}_2$  and  $\text{NO}$ ) were investigated. These results provide valuable insights into the reaction mechanisms and kinetics of imines, thereby advancing our understanding of their atmospheric chemistry and  $\bullet\text{Cl}$  chemistry.

## 2 Computational Details

### 2.1 Global Minimum Search

The selected imines can exist in multiple distinct gas-phase configurations. To identify their global minima, a multi-step conformational sampling scheme was employed, following our earlier studies (Lu et al., 2024; Ma et al., 2019; Ma et al., 2023) To explore their conformational space, the gentor module within the Molclus program was initially used to form the range of conformations (Lu, 2025). The generated conformers were further optimized at the PM6 and MP2/6-31+G(3df,2p) level (Vereecken and Francisco, 2012). Single-point energy calculations were performed at the CCSD(T)/aug-cc-pVTZ level (Vereecken and Francisco, 2012). The structures with the lowest Gibbs free energy were identified as the global minima and served as the initial structures for investigating the reaction mechanisms and kinetics. The corresponding global minima of five imines are presented in Fig. S1.

## 2.2 Electronic Structure Calculations

All electronic structure calculations were conducted with the Gaussian 09 program package (Frisch et al., 2009). For the  $\cdot\text{Cl}$ -initiated reactions of five imines, geometry optimizations and harmonic vibrational frequency calculations were performed at the MP2/6-31+G(3df,2p) level, followed by single-point energy evaluations at the CCSD(T)/aug-cc-pVTZ level. This is consistent with our previous work on  $\cdot\text{Cl}$ -initiated reactions of ONCs systems, where the combination of MP2 and CCSD(T) methods has proven to yield reliable energies (Ma et al., 2018). Considering the substantially increased computational cost for the subsequent reactions of the resulting imine radicals, we employed the M06-2X/6-31+G(d,p) level for geometry optimizations and frequency calculations (Zhao and Truhlar, 2008), followed by CCSD(T)/6-311+G(2df,2p) single-point energy evaluations (Pople et al., 1989). The M06-2X functional combined with CCSD(T) method has successfully been applied to predict radical +  $\text{O}_2$  /  $\text{NO}$  reactions (Fu et al., 2024). Intrinsic reaction coordinate (IRC) calculations were used to confirm that each transition state correctly connects the relevant reactants and products at the respective optimization levels (Fukui, 1981). The isolated  $\cdot\text{Cl}$  was treated with a correction value of  $0.8 \text{ kcal mol}^{-1}$  to account for spin-orbit coupling effects (Ma et al., 2018). Natural bond orbital (NBO) analysis was conducted to characterize the atomic charges in the transition states of initial reaction (Reed et al., 1985). Unless otherwise specified, the following labels are used throughout the manuscript: “reactants” (R), “pre-reaction complex” (RC), “post-reaction complex” (PC), “transition states” (TS), “intermediates” (IM) and “products” (P).

## 2.3 Kinetics Calculations

MultiWell 2014.1 software was employed to simulate the reaction kinetics (Barker et al., 2014; Barker, 2001). Rate constants for the multi-channel and multi-well chemical reactions involving tight transition states were calculated using Rice-Ramsberger-Kassel-Marcus (RRKM) theory (Barker, 2001), while those for barrierless reactions were obtained with long-range transition-state theory with a dispersion force potential or Inverse Laplace Transformation (ILT) model. Full computational settings are provided in our previous studies (Ma et al., 2018, 2021). For single-step processes, canonical transition state theory (TST) in the Thermo module of MultiWell-2014.1 program suite was applied to calculate the rate constants. Tunneling corrections for H-shift or H-abstraction reactions were included via a one-dimensional asymmetric Eckart potential (Eckart, 1930). Reaction rate constants for  $\cdot\text{Cl}$  + imines reactions were performed across the temperature range of 260–330 K. The parameters used in the long-range transition-state theory (LRTST) calculations and Lennard-Jones parameters for various species used in

the MultiWell are shown in Supplement Tables S1 and S2, respectively.

## 3 Results and Discussion

### 3.1 $\cdot\text{Cl}$ -Initiated Reactions

$\cdot\text{Cl}$  can either abstract H atoms from the  $=\text{CH}_x-$  ( $x = 1-2$ ) and  $=\text{NH}$  groups or add to the  $\text{C}=\text{N}$  bond of the target imines. The calculated energetic data are summarized in Table 1, with the corresponding zero-point energy (ZPE) corrected potential energy surfaces (PES) shown in Fig. S2 in the Supplement. Based on the reaction activation energies ( $E_a$ ) in Table 1, it can be concluded that H-abstractions at the N-sites is the most energetically favorable pathway for  $\cdot\text{Cl}$ -initiated reactions of  $\text{CH}_2=\text{NH}$ ,  $\text{CH}_3\text{CH}=\text{NH}$ ,  $(\text{CH}_3)_2\text{C}=\text{NH}$ , and  $\text{HN}=\text{CHCH}_2\text{OH}$ , which is similar to amines +  $\cdot\text{Cl}$  systems (Ma et al., 2018; Xue et al., 2022; Xie et al., 2015; Xie et al., 2017). Notably, the  $E_a$  values for the formation of N-centered radicals from imines are significantly higher than those from amines. Interestingly,  $\cdot\text{OH}$ -initiated reactions of  $\text{CH}_2=\text{NH}$  can also result in the formation of N-centered radical, unlike  $\cdot\text{OH}$ -initiated reactions of amines, although the  $E_a$  value for the  $\cdot\text{OH} + \text{CH}_2=\text{NH}$  reaction being approximately  $4 \text{ kcal mol}^{-1}$  higher than that of  $\cdot\text{Cl} + \text{CH}_2=\text{NH}$  reaction (Bunkan et al., 2014). For  $\text{CH}_3\text{N}=\text{CH}_2$ , H-abstraction from the  $\text{CH}_2$  site forming C-centered radical is the most favorable pathway, consistent with  $\cdot\text{OH} + \text{CH}_3\text{N}=\text{CH}_2$  system (Bunkan et al., 2022).

It should be noted that, despite numerous attempts, the TSs for H-abstraction from the  $\text{CH}_2$  site of  $\text{CH}_3\text{N}=\text{CH}_2$  and for  $\cdot\text{Cl}$  addition to the N site of  $(\text{CH}_3)_2\text{C}=\text{NH}$  could not be located at the MP2/6-31+G(3df,2p) level. By analyzing the thermodynamic data, we found that  $\cdot\text{Cl}$  addition to the N-site is thermodynamically unfavorable for the  $\cdot\text{Cl} + (\text{CH}_3)_2\text{C}=\text{NH}$  reactions, suggesting that this pathway plays a negligible role. For  $(\text{CH}_3)_2\text{C}=\text{NH}$ , the  $\text{TS}_{3-2}$  of  $\cdot\text{Cl}$  abstracting a H-atom from  $\text{CH}_2$  site was located at the MP2/6-31+G(d,p) level. To evaluate the performance of the MP2/6-31+G(d,p) method for geometry optimization and CCSD(T)/6-311+G(2df,2p) for single point energy calculation in calculating the reaction energies ( $E_a$  and  $\Delta E$ ) of H-abstraction at the  $\text{CH}_2$  site of  $\text{CH}_3\text{N}=\text{CH}_2$ , we randomly selected a pathway involved  $\text{TS}_{3-1}$  to calculate the  $E_a$  and  $\Delta E$  values using the CCSD(T)/6-311+G(2df,2p)//MP2/6-31+G(d,p) and CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p) methods. The results show that the difference in  $E_a$  and  $\Delta E$  values between the CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p) and CCSD(T)/6-311+G(2df,2p)//MP2/6-31+G(d,p) were within the quantum chemistry method ( $1.0 \text{ kcal mol}^{-1}$ ) (Table S3). These indicate that the more efficient CCSD(T)/6-311+G(2df,2p)//MP2/6-31+G(d,p) method is reliable, and the  $E_a$  values of  $\text{TS}_{3-2}$  obtained with this method have negligible effects on the conclusions.

**Table 1.** Calculated reaction activation energies ( $E_a$ , in kcal mol<sup>-1</sup>), thermodynamic energies ( $\Delta E$ , kcal mol<sup>-1</sup>), branch ratios ( $\Gamma$ ) and reaction rate constants ( $k_{Cl}$ , in cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup>) for the reactions of  $\bullet Cl$  with (a) CH<sub>2</sub>=NH, (b) CH<sub>3</sub>CH=NH, (c) CH<sub>3</sub>N=CH<sub>2</sub>, (d) (CH<sub>3</sub>)<sub>2</sub>C=NH, and (e) HN=CHCH<sub>2</sub>OH at the CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p) level of theory.

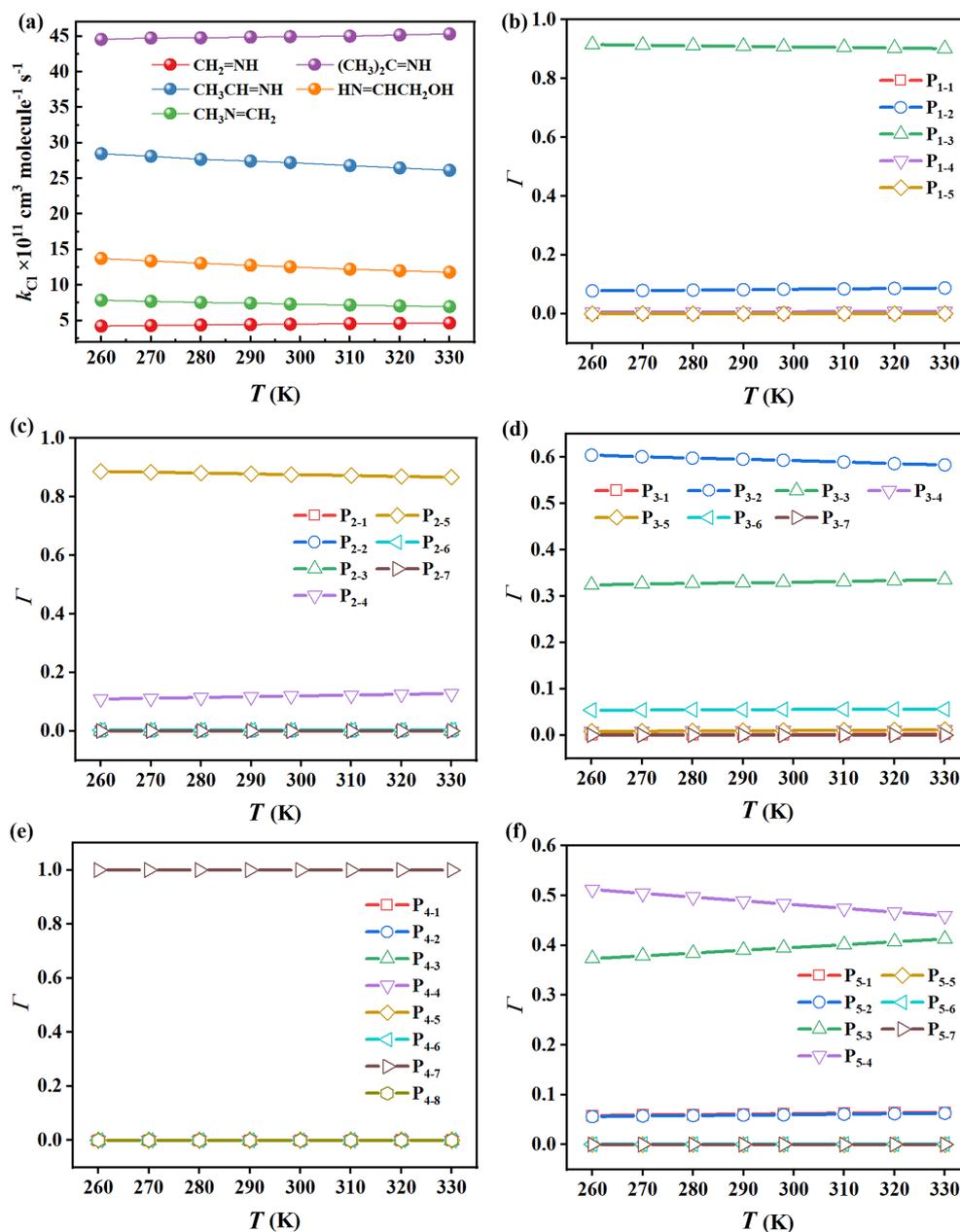
	Species	TS	$E_a$	$\Delta E$	$\Gamma$	$k_{Cl} \times 10^{-11}$
(a)	P <sub>1-1</sub>	TS <sub>1-1</sub>	2.35	0.08	0.26 %	4.5
	P <sub>1-2</sub>	TS <sub>1-2</sub>	-0.60	-4.30	8.25 %	
	P <sub>1-3</sub>	TS <sub>1-3</sub>	-4.39	-12.22	90.85 %	
	P <sub>1-4</sub>	TS <sub>1-4</sub>	0.89	-13.13	0.64 %	
	P <sub>1-5</sub>	TS <sub>1-5</sub>	5.14	8.97	0.00 %	
(b)	P <sub>2-1</sub>	TS <sub>2-1</sub>	7.95	5.68	0.00 %	27.2
	P <sub>2-2/3</sub>	TS <sub>2-2</sub>	2.26	-8.59	0.06 %	
	P <sub>2-2/3</sub>	TS <sub>2-3</sub>	2.29	-8.59	0.05 %	
	P <sub>2-4</sub>	TS <sub>2-4</sub>	-2.22	-5.15	11.96 %	
	P <sub>2-5</sub>	TS <sub>2-5</sub>	-5.99	-10.55	87.58 %	
	P <sub>2-6</sub>	TS <sub>2-6</sub>	-0.85	-11.12	0.35 %	
	P <sub>2-7</sub>	TS <sub>2-7</sub>	6.57	8.46	0.00 %	
(c)	P <sub>3-1</sub>	TS <sub>3-1</sub>	1.91	1.87	0.16 %	7.32
	P <sub>3-2</sub>	TS <sub>3-2</sub>	-2.57*	-1.66*	59.37 %	
	P <sub>3-3</sub>	TS <sub>3-3</sub>	-1.67	-2.01	32.98 %	
	P <sub>3-4/5</sub>	TS <sub>3-4/5</sub>	1.39	-12.00	0.98 %/0.98 %	
	P <sub>3-6</sub>	TS <sub>3-6</sub>	-1.22	-15.02	5.53 %	
	P <sub>3-7</sub>	TS <sub>3-7</sub>	4.26	8.19	0.00 %	
	(d)	P <sub>4-1</sub>	TS <sub>4-1</sub>	6.31	4.35	
P <sub>4-2/3</sub>		TS <sub>4-2/3</sub>	0.89	-6.82	0.01 %/0.01 %	
P <sub>4-4</sub>		TS <sub>4-4</sub>	4.19	2.94	0.00 %	
P <sub>4-5/6</sub>		TS <sub>4-5/6</sub>	1.21	-7.13	0.01 %/0.01 %	
P <sub>4-7</sub>		TS <sub>4-7</sub>	-6.87	-10.35	99.96 %	
P <sub>4-8</sub>		TS <sub>4-8</sub>	-1.40	-10.53	0.00 %	
P <sub>4-9</sub>		TS <sub>4-9</sub>	-	12.43	-	
(e)		P <sub>5-1/2</sub>	TS <sub>5-1/2</sub>	-1.13	-20.52	6.08 %/6.08 %
	P <sub>5-3</sub>	TS <sub>5-3</sub>	-1.53	-2.98	39.46 %	
	P <sub>5-4</sub>	TS <sub>5-4</sub>	-2.59	-6.96	48.33 %	
	P <sub>5-5</sub>	TS <sub>5-5</sub>	15.09	9.69	0.00 %	
	P <sub>5-6</sub>	TS <sub>5-6</sub>	1.16	-8.83	0.05 %	
	P <sub>5-7</sub>	TS <sub>5-7</sub>	9.42	12.23	0.00 %	

\* Energies were calculated at the CCSD(T)/6-311+G(2df,2p)//MP2/6-31+G(d,p) level.

Since  $\bullet Cl$  abstracts a H-atom from sp<sup>2</sup>-N site is more favorable than from other sites, it merits further discussion why the  $E_a$  values for generating N-centered radicals are substantially lower than those for C-centered radicals. By analyzing NBO charges for all TSs, we found that larger charge transfers occurred at the most favorable transition states TS<sub>1-3</sub> (-0.428  $e$ ), TS<sub>2-5</sub> (-0.447  $e$ ), TS<sub>3-2</sub> (-0.318  $e$ ) and TS<sub>4-7</sub> (-0.456  $e$ ) than other TSs (see Table S4). This indicates that charge transfer contributes critically to the stabilization of these TSs, thereby facilitating the formation of N-centered radicals in  $\bullet Cl$ -initiated reactions of NH-containing imines. A similar phenomenon was observed in the piperazine (PZ) +  $\bullet Cl$  system (Ma et al., 2018). However, in the case of  $\bullet Cl$  + HN=CHCH<sub>2</sub>OH system, the charge transfer at the TS<sub>5-5</sub> (-0.430  $e$ ) is slightly larger than that at the TS<sub>5-4</sub>

(the most favorable one, -0.423  $e$ ) (see Table S4). The presence of intermolecular hydrogen bonds in TS<sub>5-4</sub> may account for its lower  $E_a$  value (see Fig. S3).

Using the master equation approach, the overall rate constants ( $k_{Cl}$ ) were determined to be  $4.50 \times 10^{-11}$ ,  $2.72 \times 10^{-10}$ ,  $7.32 \times 10^{-11}$ ,  $4.48 \times 10^{-10}$  and  $1.26 \times 10^{-10}$  cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup> for  $\bullet Cl$  + CH<sub>2</sub>=NH,  $\bullet Cl$  + CH<sub>3</sub>CH=NH,  $\bullet Cl$  + CH<sub>3</sub>N=CH<sub>2</sub>,  $\bullet Cl$  + (CH<sub>3</sub>)<sub>2</sub>C=NH, and  $\bullet Cl$  + HN=CHCH<sub>2</sub>OH reactions at 298 K and 1 atm, respectively. The experimental  $k_{Cl}$  value available for the CH<sub>3</sub>N=CH<sub>2</sub> +  $\bullet Cl$  reaction is  $(1.9 \pm 0.15) \times 10^{-11}$  cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup> (Bunkan et al., 2022), which is in good consistency with the corresponding computational result ( $7.32 \times 10^{-11}$  cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup>). This could further support the reliability of our computa-



**Figure 1.** Reaction rate constants ( $k_{Cl}$ ) for the reactions of five imines (a), and branching ratios ( $\Gamma$ ) for products for the reactions of (b)  $\text{CH}_2=\text{NH}$ , (c)  $\text{CH}_3\text{CH}=\text{NH}$ , (d)  $\text{CH}_3\text{N}=\text{CH}_2$ , (e)  $(\text{CH}_3)_2\text{C}=\text{NH}$ , and (f)  $\text{HN}=\text{CHCH}_2\text{OH}$  initiated by  $\cdot\text{Cl}$  in the temperature range of 260–300 K at 1 atm.

tional approach. Over the temperature range of 260–330 K,  $\text{CH}_2=\text{NH}$  and  $(\text{CH}_3)_2\text{C}=\text{NH}$  show positive temperature dependence for  $k_{Cl}$ , whereas  $\text{CH}_3\text{CH}=\text{NH}$ ,  $\text{CH}_3\text{N}=\text{CH}_2$  and  $\text{HN}=\text{CHCH}_2\text{OH}$  exhibit a negative dependence (Fig. 1a). By analyzing the substitutions effects on the  $k_{Cl}$  values, it can be found that  $-\text{CH}_3$  and  $-(\text{CH}_3)_2$  substitutions at the  $>\text{CR}1=$  position, as well as the  $-\text{CH}_3$  substitution at the  $=\text{NR}2$  position, increase the  $k_{Cl}$  values, while substitutions at the  $\text{R}3\text{CCH}=\text{}$  position have little effect on the  $k_{Cl}$  values.

The calculated  $\Gamma$  values for the N-centered radicals in the reactions of  $\cdot\text{Cl} + \text{CH}_2=\text{NH}$ ,  $\cdot\text{Cl} + \text{CH}_3\text{CH}=\text{NH}$ ,  $\cdot\text{Cl} + (\text{CH}_3)_2\text{C}=\text{NH}$  and  $\cdot\text{Cl} + \text{HN}=\text{CHCH}_2\text{OH}$  are 90.85 %, 87.58 %, 99.96 % and 48.33 % at 1 atm and 298 K, respectively. The  $\Gamma$  values of N-centered radicals show a slight decrease with increasing temperature, whereas those of other product species remain very small and negligible across the studied temperature range (Fig. 1b, c, e and f). Therefore, N-centered radicals are the main products in these four reactions under atmospheric conditions, similar to

$\bullet\text{Cl}$  + amines systems (Ma et al., 2018; Xie et al., 2015). For the  $\bullet\text{Cl} + \text{CH}_3\text{N}=\text{CH}_2$  reaction, the calculated  $\Gamma$  values for  $\text{P}_{3-1}$ ,  $\text{P}_{3-2}$ ,  $\text{P}_{3-3}$ ,  $\text{P}_{3-4/5}$ ,  $\text{P}_{3-6}$ , and  $\text{P}_{3-7}$  are 0.16 %, 59.36 %, 32.98 %, 0.98 %, 5.53 %, 0.00 %, respectively, indicating a strong preference for the formation of C-centered radicals (Fig. 1d). Since previous study have investigated the transformation of  $\text{CH}_3\text{N}=\text{CH}\bullet$  ( $\text{P}_{3-2}$ ) (Bunkan et al., 2022), we mainly considered the further transformation of the formed four N-centered radicals in the subsequent sections.

### 3.2 Subsequent Reactions of the formed N-centered radicals

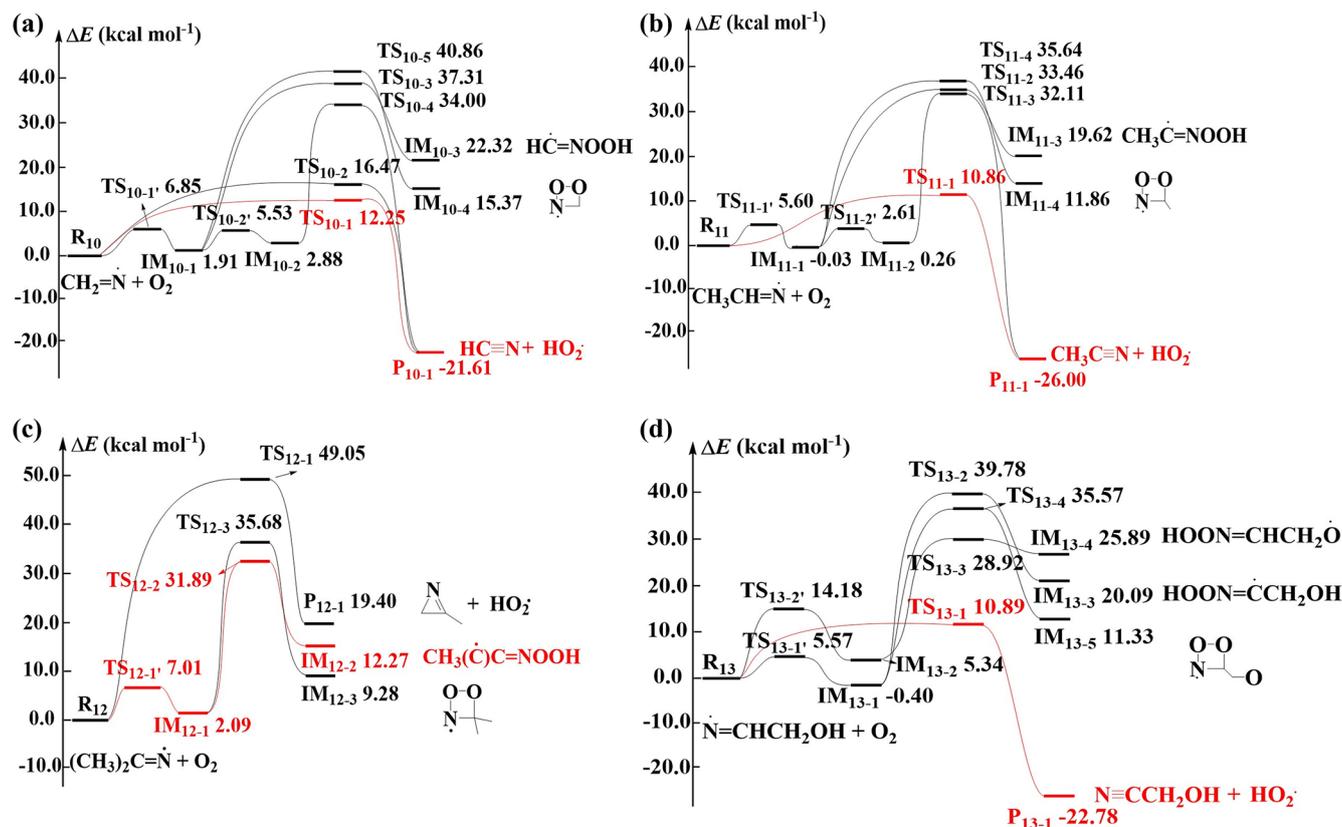
Consistent with previously studied N-centered radicals (Ma et al., 2018; Xie et al., 2015; Da Silva, 2013; Tang and Nielsen, 2012), the four N-centered radicals ( $\text{CH}_2=\text{N}\bullet$ ,  $\text{CH}_3\text{CH}=\text{N}\bullet$ ,  $(\text{CH}_3)_2\text{C}=\text{N}\bullet$ , and  $\bullet\text{N}=\text{CHCH}_2\text{OH}$ ) will subsequently undergo self-isomerization/dissociation or react with key atmospheric oxidants, such as  $\text{O}_2$  and  $\text{NO}$ . From the calculated ZPE-corrected PES of self-isomerization/dissociation for these four N-centered radicals (Fig. S4), it can be found that the  $E_a$  values are in the range of 27.85–79.58  $\text{kcal mol}^{-1}$ , suggesting that both the self-isomerization and dissociation processes proceed very slowly. Therefore, in the atmosphere, these four N-centered radicals are more likely to react with  $\text{O}_2$  and  $\text{NO}$ .

Regarding the reactions of the four N-centered radicals with  $\text{O}_2$ , two distinct routes were identified. The first pathway involves  $\text{O}_2$  directly abstracting a H-atom from  $=\text{CH}_x-$  ( $x=1-2$ ) and  $-\text{CH}_3$  sites, forming cyanide compounds, cyclic imines and  $\text{HO}_2\bullet$ . The second pathway proceeds by  $\text{O}_2$  addition to the  $=\text{N}-$  site, with  $E_a$  values in the range of 5.60–14.18  $\text{kcal mol}^{-1}$ , forming adducts with various conformations depending on the direction of  $\text{O}_2$  attack. The calculated ZPE-corrected PES for the four N-centered radicals +  $\text{O}_2$  reactions are shown in Fig. 2. With the exception of  $(\text{CH}_3)_2\text{C}=\text{N}\bullet$ , the  $E_a$  values for the direct H-abstraction pathways in the  $\text{CH}_2=\text{N}\bullet + \text{O}_2$ ,  $\text{CH}_3\text{CH}=\text{N}\bullet + \text{O}_2$ , and  $\bullet\text{N}=\text{CHCH}_2\text{OH} + \text{O}_2$  reactions are 12.25, 10.86 and 10.89  $\text{kcal mol}^{-1}$ , respectively, significantly lower than those of addition pathways. Therefore, the H-abstraction pathway represents the primary pathway for these three reactions, producing cyanide compounds and  $\text{HO}_2\bullet$ . A similar direct H-abstraction mechanism has been observed in the reactions of N-centered radicals formed from amines with  $\text{O}_2$  (Ma et al., 2018; Xue et al., 2022; Xie et al., 2015; Xie et al., 2017). For  $(\text{CH}_3)_2\text{C}=\text{N}\bullet$ , the  $E_a$  value of  $\text{O}_2$  addition forming  $\text{IM}_{12-1}$  is lower than that for the direct H-abstraction pathways. Therefore, the formation of adduct  $\text{IM}_{12-1}$  is most favorable in the initial attack of  $\text{O}_2$ . However, the  $\text{IM}_{12-1}$  is unstable and will readily reverts to  $(\text{CH}_3)_2\text{C}=\text{N}\bullet$  and  $\text{O}_2$ , making it difficult to react with  $\text{O}_2$ .

Our previous studies revealed that  $E_a$  values of  $\text{O}_2$  directly abstracting H-atoms from N-centered radicals is highly sensitive to the employed theoretical methods, which

ultimately affects the reaction rate constant ( $k_{\text{O}_2}$ ) with  $\text{O}_2$  (Liu et al., 2019; Xie et al., 2017). Since  $k_{\text{O}_2}$  is a key parameter in determining nitrosamines yields in subsequent reactions (Liu et al., 2019; Wang, 2015), we further evaluated the impact of different computational approaches on the  $E_a$  values for direct H-abstraction pathways. The  $E_a$  values obtained at  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$ ,  $\text{CCSD(T)/aug-cc-pVTZ//MP2/aug-cc-pVTZ}$  and  $\text{CCSD(T)/6-311+G(2df,2p)//M06-2X/6-31+G(d,p)}$  are summarized in Table S5. When the geometry optimization method was changed from  $\text{M06-2X/6-31+G(d,p)}$  to  $\text{MP2/6-31+G(3df,2p)}$  or  $\text{MP2/aug-cc-pVTZ}$ , and the single point energy calculation was changed from  $\text{CCSD(T)/6-311+G(2df,2p)}$  to  $\text{CCSD(T)/aug-cc-pVTZ}$ , the deviations in  $E_a$  values reach up to 6.2  $\text{kcal mol}^{-1}$  [between  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$  and  $\text{CCSD(T)/6-311+G(2df,2p)//M06-2X/6-31+G(d,p)}$ ] and 6.4  $\text{kcal mol}^{-1}$  [between  $\text{CCSD(T)/aug-cc-pVTZ//MP2/aug-cc-pVTZ}$  and  $\text{CCSD(T)/6-311+G(2df,2p)//M06-2X/6-31+G(d,p)}$ ]. By contrast, changing the optimization method from  $\text{MP2/6-31+G(3df,2p)}$  to  $\text{MP2/aug-cc-pVTZ}$  lead to only 0.3  $\text{kcal mol}^{-1}$  deviation of  $E_a$  values between  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$  and  $\text{CCSD(T)/aug-cc-pVTZ//MP2/aug-cc-pVTZ}$ . These results indicate that the combination of the computationally cheaper  $\text{M06-2X}$  and  $\text{CCSD(T)}$  does not provide accurate results for these systems. Therefore, the  $E_a$  values calculated at the  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$  level were adopted for subsequent kinetics calculations. Notably, H-abstraction in the  $\text{CH}_2=\text{N}\bullet + \text{O}_2$ ,  $\text{CH}_3\text{CH}=\text{N}\bullet + \text{O}_2$ , and  $\bullet\text{N}=\text{CHCH}_2\text{OH} + \text{O}_2$  systems remain the most favorable reaction pathways even at higher  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$  methods.

Since  $E_a$  values of the direct H-abstraction pathways are notably lower than those of the addition pathways for  $\text{CH}_2=\text{N}\bullet$ ,  $\text{CH}_3\text{CH}=\text{N}\bullet$  and  $\bullet\text{N}=\text{CHCH}_2\text{OH}$ , the overall  $k_{\text{O}_2}$  are approximated to be the rate constants of the direct H-abstraction pathways. The TST was applied to examine the kinetics of the direct H-abstraction pathways. The calculated  $k_{\text{O}_2}$  of the direct H-abstraction pathway for  $\text{CH}_2=\text{N}\bullet + \text{O}_2$ ,  $\text{CH}_3\text{CH}=\text{N}\bullet + \text{O}_2$  and  $\bullet\text{N}=\text{CHCH}_2\text{OH} + \text{O}_2$  reactions based on the energies calculated at the  $\text{CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p)}$  level are  $8.94 \times 10^{-18}$ ,  $1.15 \times 10^{-17}$  and  $1.38 \times 10^{-17} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$  at 298 K, respectively. The  $k_{\text{O}_2}$  values of these three N-centered radicals formed from imines oxidation are similar to or higher than those of chain- and cyclic-like N-centered radicals from amines oxidation (Ma et al., 2018; Liu et al., 2019; Xie et al., 2015). Therefore, the  $k_{\text{O}_2}$  values for N-centered radicals with  $\text{O}_2$  varies greatly with their specific molecular structures. Previous studies have shown that the reactions of amines-derived N-centered radicals with  $\text{O}_2$  can compete with their reactions with  $\text{NO}$  under typical tropospheric  $\text{NO}$  concentrations (Ma et al., 2018). Therefore, we further investigated the reactions



**Figure 2.** Schematic ZPE-corrected PES for the reactions of O<sub>2</sub> with (a) CH<sub>2</sub>=N•, (b) CH<sub>3</sub>CH=N•, (c) (CH<sub>3</sub>)<sub>2</sub>C=N•, and (d) •N=CHCH<sub>2</sub>OH at the CCSD(T)/6-311+G(2df,2p)//M06-2X/6-31+G(d,p) level of theory.

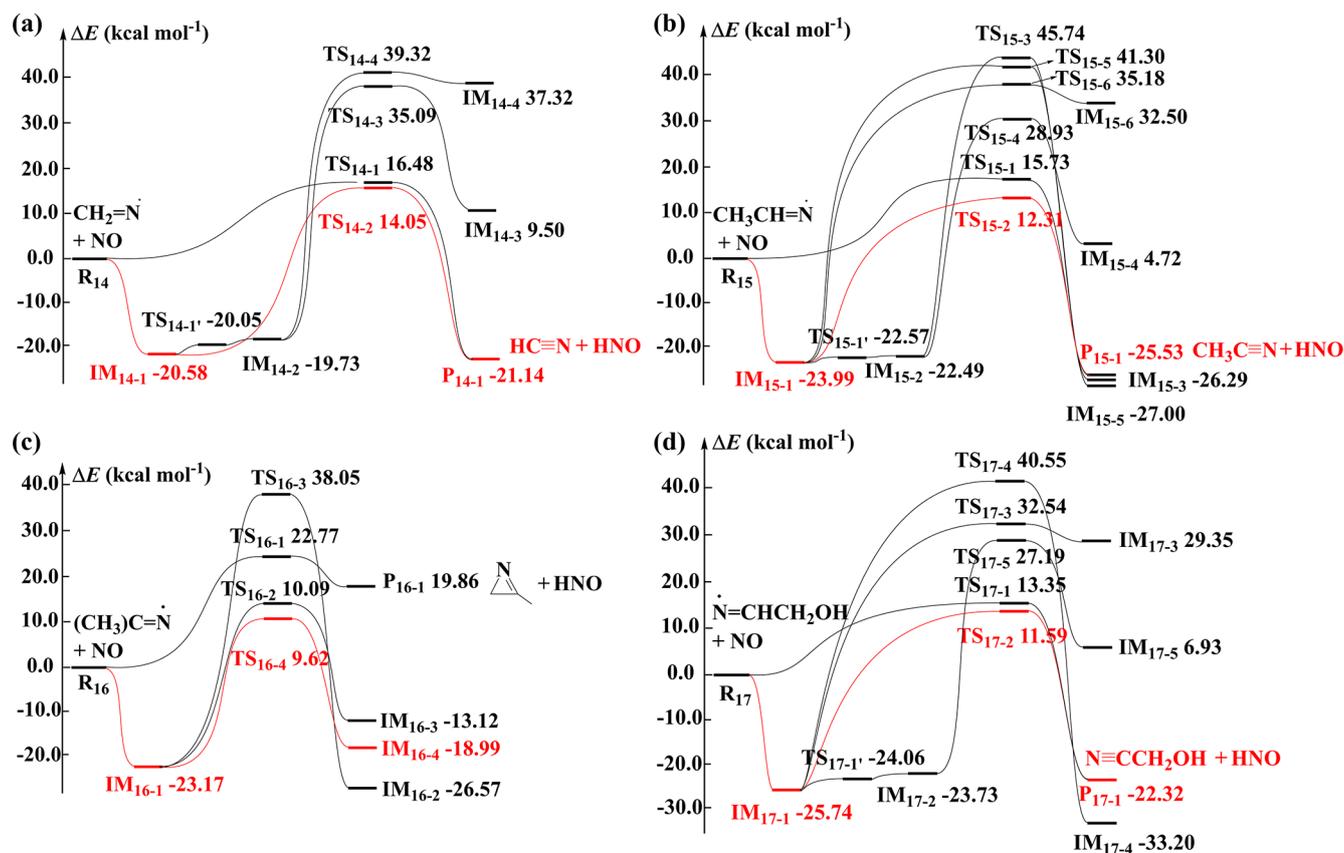
of these four N-centered radicals (CH<sub>2</sub>=N•, CH<sub>3</sub>CH=N•, (CH<sub>3</sub>)<sub>2</sub>C=N• and •N=CHCH<sub>2</sub>OH) with NO.

The calculated ZPE-corrected PES for the reactions of these four N-centered radicals with NO are presented in Fig. 3. As shown in Fig. 3, two types of reaction pathways are observed during the initial interaction of NO with the N-centered radicals. The first is the direct H-abstraction pathway, where NO abstracts a H-atom from the =CH<sub>x</sub>- (x = 1–2) sites adjacent to the =N- and -CH<sub>3</sub> groups. These H-abstraction pathways need to overcome at least 13.35 kcal mol<sup>-1</sup> *E<sub>a</sub>* values to form cyanide compounds and HNO. The second is the NO addition pathway, where NO barrierlessly addition to the =N- site of the four N-centered radicals, forming nitrosamines adducts with different conformations depending on the approach direction of NO. Interconversion between these adducts, such as IM<sub>14-1</sub> and IM<sub>14-2</sub>, IM<sub>15-1</sub> and IM<sub>15-2</sub>, IM<sub>17-1</sub> and IM<sub>17-2</sub> proceeds with very low energy barriers of 0.53, 1.42 and 1.68 kcal mol<sup>-1</sup>, respectively.

The formed four adducts (nitrosamines) can further undergo isomerization or dissociation reactions. For IM<sub>14-1</sub>/IM<sub>14-2</sub>, IM<sub>15-1</sub>/IM<sub>15-2</sub>, IM<sub>16-1</sub>, IM<sub>17-1</sub>/IM<sub>17-2</sub>, three, five, three and four H-shift pathways are identified, respectively. It is observed that H-shifts from the =CH<sub>x</sub>- (x = 1–2)

sites adjacent to the =N- and -CH<sub>3</sub> groups to the O-atom of the -NNO group are the most favorable. However, the formed four adducts need to overcome high energy barriers to isomerize or dissociate into fragmentation products. This mechanism is analogous to amines-derived N-centered radicals with NO. In addition, the main reaction pathway remains even at high computational methods (see Table S6 and S7).

To maintain consistency with the N-centered radicals + O<sub>2</sub> reactions, the reaction energies calculated at the CCSD(T)/aug-cc-pVTZ//MP2/6-31+G(3df,2p) level are used to calculate the reaction rate constant (*k*<sub>NO</sub>) for the reactions of four N-centered radicals with NO. The calculated reaction rate constants for NO addition pathways for the reactions of CH<sub>2</sub>=N• + NO, CH<sub>3</sub>CH=N• + NO, (CH<sub>3</sub>)<sub>2</sub>C=N• + NO and •N=CHCH<sub>2</sub>OH + NO are 2.20 × 10<sup>-16</sup>, 1.42 × 10<sup>-12</sup>, 1.09 × 10<sup>-10</sup> and 4.30 × 10<sup>-11</sup> cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup> at 298 K, respectively. These are much higher than those for the corresponding direct H-abstraction pathways (1.85 × 10<sup>-21</sup>, 2.87 × 10<sup>-21</sup>, 5.57 × 10<sup>-27</sup> and 1.04 × 10<sup>-19</sup> cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup>). Therefore, the *k*<sub>NO</sub> can be assumed to be equal to the reaction rate constant for the addition pathways. The calculated  $\Gamma$  values for the nitrosamine (IM<sub>14-1</sub>, IM<sub>15-1</sub>, IM<sub>16-1</sub> and IM<sub>17-1</sub>) for the reactions of CH<sub>2</sub>=N• + NO, CH<sub>3</sub>CH=N• + NO,



**Figure 3.** Schematic ZPE-corrected PES for the reactions of NO with (a)  $\text{CH}_2=\text{N}\cdot$ , (b)  $\text{CH}_3\text{CH}=\text{N}\cdot$ , (c)  $(\text{CH}_3)_2\text{C}=\text{N}\cdot$ , and (d)  $\cdot\text{N}=\text{CHCH}_2\text{OH}$  at the CCSD(T)/6-311+G(2df,2p)//M06-2X/6-31+G(d,p) level of theory.

$(\text{CH}_3)_2\text{C}=\text{N}\cdot + \text{NO}$  and  $\cdot\text{N}=\text{CHCH}_2\text{OH} + \text{NO}$  are 0.08 %, 6.22 %, 52.56 % and 35.32 % at 1 atm and 298 K, respectively. Consequently, except for  $\text{CH}_2=\text{N}\cdot$  and  $\text{CH}_3\text{CH}=\text{N}\cdot$ , the reactions of  $(\text{CH}_3)_2\text{C}=\text{N}\cdot$  and  $\cdot\text{N}=\text{CHCH}_2\text{OH}$  with NO can lead to carcinogenic nitrosamines with high yields. It is noteworthy that the  $\Gamma$  values for nitrosamine formation in these two reactions are significantly lower than those of PZ-N (99.97 %) (Ma et al., 2018), Monoethanolamine (MEA)-N (86 %) (Xie et al., 2015) and  $\text{CH}_3\text{NH}\cdot$  ( $\sim 60\%$ ) (Da Silva, 2013) with NO.

Since  $(\text{CH}_3)_2\text{C}=\text{N}\cdot$  exclusively reaction with NO, we only assessed the competition of  $\text{O}_2$  and NO for other three N-centered radicals using the calculated  $k_{\text{O}_2}$  and  $k_{\text{NO}}$  values. The required concentrations of NO ( $[\text{NO}]$ ) to equalize the pseudo-first-order rate constants for N-centered radicals with  $\text{O}_2$  are  $8.1 \times 10^6$ ,  $1.7 \times 10^3$  and 64 ppb, respectively. For  $\text{CH}_2=\text{N}\cdot$  and  $\text{CH}_3\text{CH}=\text{N}\cdot$ , the required  $[\text{NO}]$  are very high, far exceeding the typical  $[\text{NO}]$  encountered in the atmosphere (Ren, 2003). Therefore,  $\text{CH}_2=\text{N}\cdot$  and  $\text{CH}_3\text{CH}=\text{N}\cdot$  are expected to primarily react with  $\text{O}_2$  to form  $\text{HC}\equiv\text{N}$  and  $\text{CH}_3\text{C}\equiv\text{N}$ , respectively, consistent with previous studies of  $\text{CH}_2=\text{NH}$  (Bunkan et al., 2014). In contrast, for  $\cdot\text{N}=\text{CHCH}_2\text{OH}$ , the required  $[\text{NO}]$  are achievable under polluted atmospheric conditions. Therefore, both  $(\text{CH}_3)_2\text{C}=\text{N}\cdot$

and  $\cdot\text{N}=\text{CHCH}_2\text{OH}$  can react with NO to form nitrosamines. To the best of knowledge, this study represents the first report demonstrating that N-centered radicals formed from  $\cdot\text{Cl}$ -initiated reactions of imines can lead to carcinogenic nitrosamine formation. However, the yields of nitrosamine formation are highly dependent on the specific structures of N-centered radicals.

## 4 Conclusions

Similar to previous studies on amines +  $\cdot\text{Cl}$  reaction systems, the reaction of imines containing  $=\text{NH}$  groups ( $\text{CH}_2=\text{NH}$ ,  $\text{CH}_3\text{CH}=\text{NH}$ ,  $(\text{CH}_3)_2\text{C}=\text{NH}$ , and  $\text{HN}=\text{CHCH}_2\text{OH}$ ) with  $\cdot\text{Cl}$  predominantly yield N-centered radicals. The calculated  $k_{\text{Cl}}$  values for the reactions of  $\text{CH}_2=\text{NH}$ ,  $\text{CH}_3\text{CH}=\text{NH}$ ,  $\text{CH}_3\text{N}=\text{CH}_2$ ,  $(\text{CH}_3)_2\text{C}=\text{NH}$ , and  $\text{HN}=\text{CHCH}_2\text{OH}$  are  $4.50 \times 10^{-11}$ ,  $2.72 \times 10^{-10}$ ,  $7.32 \times 10^{-11}$ ,  $4.48 \times 10^{-10}$  and  $1.26 \times 10^{-10} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$  at 298 K and 1 atm, respectively. Among the five imines studied, only the reaction kinetics of  $\text{CH}_2=\text{NH}$  and  $\text{CH}_3\text{N}=\text{CH}_2$  with  $\cdot\text{OH}$  are available ( $\text{CH}_2=\text{NH}$ :  $3.00 \times 10^{-12} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ ,  $\text{CH}_3\text{N}=\text{CH}_2$ :  $3.70 \times 10^{-12} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ ). With these data, we can estimate the contribution of  $\cdot\text{Cl}$  to the degra-

dation of  $\text{CH}_2=\text{NH}$  and  $\text{CH}_3\text{N}=\text{CH}_2$  based on the  $\bullet\text{Cl}$  concentration ( $[\bullet\text{Cl}]$ ). In the marine boundary layer, the  $[\bullet\text{Cl}]$  is typically 1%–10% of the  $[\bullet\text{OH}]$ . The contribution of  $\bullet\text{Cl}$  relative to  $\bullet\text{OH}$  ( $k_{\text{Cl}}[\bullet\text{Cl}]/k_{\text{OH}}[\bullet\text{OH}]$ ) to the transformation of  $\text{CH}_2=\text{NH}$  and  $\text{CH}_3\text{N}=\text{CH}_2$  are estimated to be 15%–150% and 20%–200%, respectively. Furthermore, the contribution of  $\bullet\text{Cl}$  relative to  $\bullet\text{OH}$  to the formation of  $\text{CH}_2=\text{N}\bullet$  is estimated at 34%–340% (estimated by  $k_{\text{Cl}}[\bullet\text{Cl}] \times \Gamma_{\text{N,Cl}}/k_{\text{OH}}[\bullet\text{OH}] \times \Gamma_{\text{N,OH}}$ , where  $\Gamma_{\text{N,Cl}}$  and  $\Gamma_{\text{N,OH}}$  represent the yields of N-centered radicals from the reactions initiated by  $\bullet\text{Cl}$  and  $\bullet\text{OH}$ , respectively). This clearly demonstrates the significant role of  $\bullet\text{Cl}$  in the transformation of  $\text{CH}_2=\text{NH}$  and  $\text{CH}_3\text{N}=\text{CH}_2$ . Although complete data on the reactions of other imines ( $\text{CH}_3\text{CH}=\text{NH}$ ,  $(\text{CH}_3)_2\text{C}=\text{NH}$ , and  $\text{HN}=\text{CHCH}_2\text{OH}$ ) with  $\bullet\text{OH}$  are lacking, it is reasonable to believe that  $\bullet\text{Cl}$  also plays a crucial role in their transformation based on their high  $k_{\text{Cl}}$  values.

Unlike N-centered radicals generated from amines oxidation, this study reveals that both  $k_{\text{O}_2}$  and  $k_{\text{NO}}$  values for N-centered radicals generated from imines oxidation are strongly dependent on their specific structures, ultimately affecting nitrosamine formation. This study provides the first evidence that N-centered radicals formed from imines oxidation can yield nitrosamines under polluted atmospheric conditions. Therefore, to comprehensively evaluate the formation of nitrosamines from imines, further investigations into the reactions of imine-derived N-centered radicals with  $\text{O}_2$  and  $\text{NO}$  are warranted.

**Data availability.** All data are available in the main text or Supplement. The other relevant data are available upon request from the corresponding authors.

**Supplement.** The supplement related to this article is available online at <https://doi.org/10.5194/acp-26-3211-2026-supplement>.

**Author contributions.** HBX and MFF designed research; XQ, MFF and HBX performed research; XQ, MFF, LC, ZQJ and HBX analyzed data; XQ and MFF wrote the paper; XQ, MFF, CJW and HBX reviewed and revised the paper.

**Competing interests.** The contact author has declared that none of the authors has any competing interests.

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**Acknowledgements.** This study was supported by Hefei Advanced Computing Center.

**Financial support.** This research has been supported by the National Natural Science Foundation of China (grant nos. 22206020, 22236004, and 22176022) and the National Key Research and Development Program of China (grant no. 2022YFC3701000).

**Review statement.** This paper was edited by Qiang Zhang and reviewed by two anonymous referees.

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