



Investigations into the gas-phase photolysis and OH radical kinetics of nitrocatechols: implications of intramolecular interactions on their atmospheric behaviour

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Abstract. The Environmental Simulation Chamber made of Quartz from the University “Alexandru Ioan Cuza” (ESC-Q-UAIC), at Iasi, Romania, was used to investigate the gas-phase reaction rate coefficients for four nitrocatechols toward OH radicals under simulated atmospheric conditions. Employing relative rate techniques at a temperature of 298 ± 2 K and a total air pressure of 1 atm, the obtained rate coefficients (in 10^{-12} cm³ s⁻¹) were as follows: $k_{3\text{NCAT}} = (3.41 \pm 0.37)$ for 3-nitrocatechol and $k_{5\text{M3NCAT}} = (5.55 \pm 0.45)$ for 5-methyl-3-nitrocatechol at 365 nm, using CH₃ONO photolysis as OH radicals source and dimethyl ether and cyclohexane as reference compounds, and $k_{4\text{NCAT}} = (1.27 \pm 0.19)$ for 4-nitrocatechol and $k_{4\text{M5NCAT}} = (0.92 \pm 0.14)$ for 4-methyl-5-nitrocatechol at 254 nm using H₂O₂ as OH radicals source and dimethyl ether and methanol as reference compounds. The photolysis rates in the actinic region, scaled to atmospheric relevant conditions by NO₂ photolysis, were evaluated for 3-nitrocatechol and 5-methyl-3-nitrocatechol: $J_{3\text{NCAT}} = (3.06 \pm 0.16) \times 10^{-4}$ s⁻¹ and $J_{5\text{M3NCAT}} = (2.14 \pm 0.18) \times 10^{-4}$ s⁻¹, respectively. The photolysis rate constants at 254 nm were measured for 4-nitrocatechol and 4-methyl-5-nitrocatechol and the obtained values are $J_{4\text{NCAT}} = (6.7 \pm 0.1) \times 10^{-5}$ s⁻¹ and $J_{4\text{M5NCAT}} = (3.2 \pm 0.3) \times 10^{-5}$ s⁻¹. Considering the obtained results, our study suggests that photolysis may be the main degradation process for 3-nitrocatechol and 5-methyl-3-nitrocatechol in the atmosphere, with a photolytic lifetime in the atmosphere of up to 2 h. Results are discussed in terms of the reactivity of the four nitrocatechols under investigation toward OH-radical-initiated oxidation and their structural features. The rate coefficient values of the nitrocatechols are also compared with those estimated from the structure-activity relationship for monocyclic aromatic hydrocarbons and assessed in relation to their gas-phase IR spectra. Additional comparison with similar compounds is also presented, underlining the implications toward possible degradation pathways and atmospheric behaviour.

1 Introduction

Aromatic hydrocarbons (AHs) are a class of volatile organic compounds (VOCs) present as primary pollutants in the atmosphere mainly because of anthropogenic activities. In urban areas, AHs are present owing to the use of solvents, incomplete combustion of fossil fuels, car-engine emissions

and industrial processes (Piccot et al., 1992). Various monocyclic aromatics occur from biomass burning in rural and remote areas (Schauer et al., 2001). Atmospheric removal of these compounds occurs via reactions with different oxidants (Finlayson-Pitts and Pitts, 2000) and by wet and dry deposition (Warneck, 2000; Calvert et al., 2002). The gas-phase oxidation initiated by the OH radicals is the main reaction re-

sponsible for the atmospheric sink of aromatic hydrocarbons. The chemical degradation of AHs forms oxidation products that could further be aromatic in nature or that could lead to ring opening and fragmentation. AHs are also known to be important contributors to photo-oxidant and secondary organic aerosol (SOA) formation in the atmosphere (Calvert et al., 2002; Jenkin et al., 2017).

Hydroxylated aromatic compounds, such as phenol (Atkinson et al., 1992; Sørensen et al., 2002; Berndt and Böge, 2003), cresols (Atkinson et al., 1992; Coeur-Tourneur et al., 2006), dimethylphenols (Thüner et al., 2004), trimethylphenols (Bejan et al., 2012), methoxyphenols (Lauraguais et al., 2015), and catechols (Olariu et al., 2000) are the most reactive species toward OH radicals. According to the Master Chemical Mechanism (MCM), version 3.3.1 (Bloss et al., 2005), which describes the detailed gas-phase chemical processes involved in the tropospheric degradation of a series of primary emitted VOCs, the major oxidation pathway for monocyclic aromatic compounds is the addition of the OH radical to the aromatic ring. This may lead to ring-retaining products. The distribution of the hydroxylated isomers varies for each monocyclic aromatic compound depending on the stability of its OH adduct. Thus, the OH-addition reactivity channel leads to the formation of phenol and cresols with a range of yields from 5 % to 50 % in the case of benzene and toluene. These formation yields are very much affected by the levels of NO_x ($\text{NO}_x = \text{NO} + \text{NO}_2$) in the reaction mixture, and the formation yields of the mono-hydroxylated aromatic compounds is decreasing in the presence of high NO_x concentration (Atkinson and Aschmann, 1994; Klotz et al., 1998, 2002; Volkamer et al., 2002). The hydroxylated products, when photo-oxidized by the OH radicals under atmospheric conditions, in turn form catechols with formation yields of about 65 % to 90 % (Olariu et al., 2002). Formation of nitrophenols and nitrocatechols from the oxidation of phenols and catechols initiated by the OH radicals is assigned to the H-atom abstraction pathway from the phenolic hydroxyl group, accounting for about 10 % in the mono-hydroxylated aromatic compounds and for about 30 % in the di-hydroxylated aromatic compounds (Atkinson et al., 1992; Olariu et al., 2002; Finewax et al., 2018). High formation yields of nitrophenols (more than 50 %) and nitrocatechols (close to 100 %) were reported from the gas-phase reactions of the NO_3 radical with the phenols and catechols, which may support the fact that the NO_3 chemistry may play an important role in the formation of such nitroaromatic compounds (Olariu et al., 2004, 2013; Finewax et al., 2018).

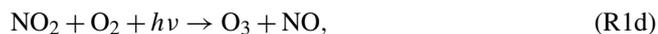
Nitroaromatic hydrocarbon compounds (NAHs) have been found in particulate matter (Herterich and Herrmann, 1990; Delhomme et al., 2010), rainwater (Grosjean, 1991; Belloli et al., 2006), clouds (Lüttke et al., 1997), water (Herterich, 1991; Rubio et al., 2012), snow (Leuenerberger et al., 1988), soil (Vozňáková et al., 1996), and fog (Richartz et al., 1990; Harrison et al., 2005), and even in remote areas such as in Terra Nova Bay, Antarctica (Vanni et al.,

2001). More recently, nitroaromatics have been reported in aerosol composition (Vidović et al., 2018) from urban area samples from Nagoya, Japan (Ikemori et al., 2019), Paris, France (Lanzafame et al., 2021), and Beijing, China (Wang et al., 2021). Nitrocatechols have been found as tracer compounds in particulate matter emitted during biomass burning, correlating well with the levoglucosan and NO_x levels (Kitanovski et al., 2012b, a; Salvador et al., 2021). In organic aerosols, the 4-nitrocatechol concentration exceeds in some cases 150 ng m^{-3} , whereas for methylated nitrocatechols the overall concentration can be as high as 821 ng m^{-3} (Inuma et al., 2010; Kitanovski et al., 2021). These high concentrations of nitrated monoaromatic hydrocarbons were determined in the winter season and have been correlated with household wood-burning use. However, in these studies, all identified nitrocatechols may cover up to 96 % of the nitrated monoaromatic hydrocarbons found in PM_{10} samples (Kitanovski et al., 2021). 4-nitrocatechol was found in the reaction of guaiacol initiated by OH radicals in the presence of NO_x (Lauraguais et al., 2014). Besides 4-nitro-, 3-nitro-, and 6-nitroguaiacol formed as oxidation products, the occurrence of 4-nitrocatechol was explained via the *-ipso* addition channel of OH radical to the methoxy group. Schwantes et al. (2017) report the formation of highly oxygenated, low-volatile products from the chamber studies of *o*-cresol oxidation under high NO_x conditions. The signal with $m/z = 169$ detected by mass spectrometric techniques has been attributed to methyl-dihydroxy-nitrobenzene. Finewax et al. (2018) have reported 4-nitrocatechol in secondary organic aerosol (SOA) composition from the gas-phase oxidation of catechol initiated by both OH (30 %) and NO_3 (91 %) radicals, respectively. Atmospheric production of nitroaromatic compounds, including nitrophenols, dinitrophenols, and nitrocatechols, is of great interest as they play a significant role in the formation of brown carbon and aerosols (Lin et al., 2016). Their low atmospheric reactivity and great ability to absorb large amounts of near-UV, visible and infrared radiation could lead to positive radiative forcing (Bejan et al., 2007; Zhang et al., 2017).

The present study is aimed at determining and discussing, to our knowledge for the first time, the OH kinetic rate coefficients for four nitrocatechols: 3-nitrocatechol (3NCAT), 4-nitrocatechol (4-NCAT), 5-methyl-3-nitrocatechol (5M3NCAT), and 4-methyl-5-nitrocatechol (4M5NCAT) by using the Environmental Simulation Chamber made of Quartz from University “Alexandru Ioan Cuza” at Iasi, Romania (ESC-Q-UAIC). Based on the ESC-Q-UAIC facilities, the photolysis frequencies of the investigated compounds were evaluated and scaled to the relevant atmospheric conditions. Detailed investigations with impact on the reactivity of highly substituted nitro-aromatic compounds are also depicted in this work.

2 Experimental

Gas-phase rate coefficients of the studied nitrocatechols toward OH radicals have been determined using the relative kinetic method. For this purpose, the ESC-Q-UAIC reactor, a 760-L and 4.2-m-long cylindrical photoreactor consisting of three quartz tubes surrounded by 32 Philips TL-DK 36W (with $\lambda_{\text{max}} = 365$ nm) and 32 UV-C TUV 30W/G30 T8 (with $\lambda_{\text{max}} = 254$ nm), has been employed. Experiments were performed at 1 atm of air and a temperature of 298 ± 2 K. Decay of the nitrocatechols and reference compounds in the reactor vessel was monitored by IR spectroscopy using a Bruker Vertex 80 FTIR coupled to a White type mirrors system mounted inside the chamber, providing a total optical path of 492 ± 1 m. Solid nitrocatechols were transferred into the reactor at low pressure, via a preheated glassware port. A stream of nitrogen flowing over the heated nitroaromatic compound was used to carry the reactants and reference compounds inside the reaction chamber. Two Teflon-blade fans were used to ensure homogeneous mixture. Dimethyl ether ($\text{C}_2\text{H}_6\text{O}$), cyclohexane (C_6H_{12}) and methanol (CH_3OH) were chosen as reference compounds in the relative kinetic study owing to their IR spectral features and well-known rate coefficients for the reaction with OH radicals, rate coefficients close to those of the investigated compounds: $k_{\text{C}_2\text{H}_6\text{O}} = (2.80 \pm 0.56) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ (Atkinson et al., 2004), $k_{\text{C}_6\text{H}_{12}} = (6.38 \pm 0.56) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ (Wilson et al., 2006) and $k_{\text{CH}_3\text{OH}} = (0.90 \pm 0.18) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ (Atkinson et al., 2004). Dimethyl ether (DME) was used as reference hydrocarbon for all the investigated nitrocatechols whereas cyclohexane was used for the kinetic studies involving 3-nitrocatechol and 5-methyl-3-nitrocatechol. Methanol was employed for the kinetic experiments of 4-nitrocatechol and 4-methyl-5-nitrocatechol. The OH radicals were generated in situ, (1) via photolysis of a mixture of $\text{CH}_3\text{ONO}/\text{NO}$ at 365 nm as shown in Reactions (R1a) to (R1e), where NO has been added in excess to suppress ozone formation and potential interference processes in the reaction kinetics:

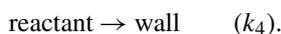
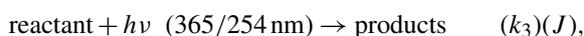
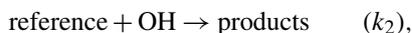
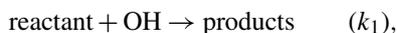


And (2) via photolysis of H_2O_2 at 254 nm for 4-nitrocatechol and 4-methyl-5-nitrocatechol as presented in Reaction (R2):



Preliminary tests showed that nitrocatechols photodissociate under experimental conditions. Corrections to reactants' concentration are needed to be made owing to the photolysis and

wall deposition, as presented in the following reactions sequence:



By representing the total decay of both reference and reactant in time with respect to the correction mentioned above, the following kinetic expression shown in Eq. (1) could be used to determine the nitroaromatic rate coefficients:

$$\ln \frac{[\text{reactant}]_{t_0}}{[\text{reactant}]_t} - (k_3 + k_4)(t - t_0) = \frac{k_1}{k_2} \ln \frac{[\text{reference}]_{t_0}}{[\text{reference}]_t}, \quad (1)$$

where $[\text{reactant}]_{t_0}$ and $[\text{reference}]_{t_0}$ are the initial concentrations of nitrocatechols and the reference compound at t_0 , and $[\text{reactant}]_t$ and $[\text{reference}]_t$ are the concentrations of nitrocatechols and the reference compounds during the reaction time t , k_1 , k_2 are the gas-phase rate coefficients for the OH radical reactions with the reactants and reference compounds respectively, k_3 is the photolysis frequency (J) of the reactants at 365/254 nm and k_4 is the wall loss rate coefficient. The ratio of k_1/k_2 determined from the slope of the linear regression is used further to obtain the rate coefficient k_1 by using well-determined k_2 rate coefficients of the reference compounds. No wall loss or photolysis was observed for the reference hydrocarbons used in this study.

However, the photolysis of nitrocatechols at 254 and 365 nm has been measured using CO as OH-radical scavenger. Photolysis rate coefficients are corrected for wall loss and used to correct further rate coefficients from the reactions of nitrocatechols with OH radicals.

Initial concentrations of reactants in the reactor were: $6.5 \times 10^{13} \text{ cm}^{-3}$ for DME, $4.8 \times 10^{13} \text{ cm}^{-3}$ for cyclohexane, $12.5 \times 10^{13} \text{ cm}^{-3}$ for methanol, $12.9 \times 10^{13} \text{ cm}^{-3}$ for NO, $16.2 \times 10^{13} \text{ cm}^{-3}$ for CH_3ONO , a maximum of $25.2 \times 10^{13} \text{ cm}^{-3}$ for H_2O_2 . For the investigated nitrocatechols, their initial concentration varied between (2.5 and $7.7) \times 10^{13} \text{ cm}^{-3}$.

3 Chemicals

Chemical compounds used in the present study were: DME > 99.9 % (suitable for GC analysis, Sigma-Aldrich), cyclohexane > 99.5 % (anhydrous, Sigma-Aldrich), methanol (anhydrous, > 99.9 % suitable for HPLC analysis, Sigma-Aldrich), NO (> 99.5 % purity, Linde), H_2O_2 (40 % solution in water, Sigma-Aldrich), CO (> 99.997 % purity, Linde). CH_3ONO was prepared in our laboratory using an adapted method from Taylor et al. (1980). All nitrocatechols have been synthesized using two methods of nitration according to available literature information proposed by Rosenblatt et al. (1953) and adapted by Iinuma et al. (2010)

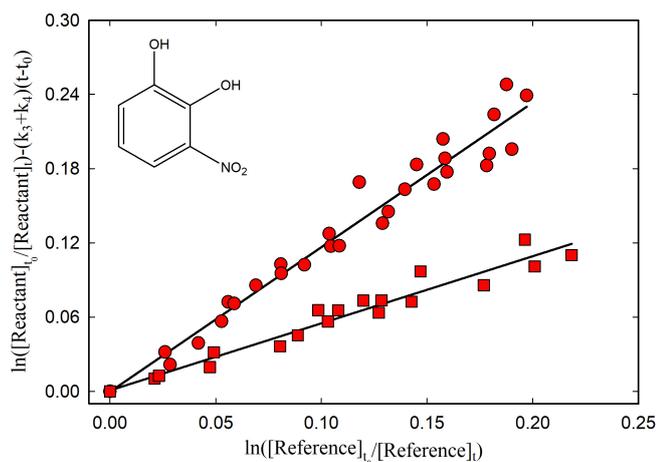


Figure 1. Kinetic plot according to Eq. (1) for the reaction of 3-nitrocatechol with OH radicals relative to dimethyl ether (red circle) and cyclohexane (red square), using photolysis of $\text{CH}_3\text{ONO}/\text{NO}$ mixture at 365 nm as OH-radical source, with $(k_3 + k_4) = (4.7 \pm 0.41) \times 10^{-4} \text{ s}^{-1}$.

and Palumbo et al. (2002). Thus, both 3-nitrocatechols have been obtained by treating an anhydrous diethyl ether solution of the corresponding catechols with freshly prepared fuming nitric acid. 4-nitrocatechol and 4-methyl-5-nitrocatechol were prepared by treating a cooled aqueous solution of catechol and sodium nitrite with sulfuric acid. Extractions and purifications of the synthesized compounds were conducted by means of sublimation and recrystallization. The purity of the synthesized compounds from H-NMR analysis was found to be about 98 % (Roman et al., 2022).

4 Results

Kinetic results for the reaction of nitrocatechols with OH radicals are presented in Figs. 1 to 4. For each compound, a minimum of four experiments have been performed. Despite difficulties in handling these low volatile compounds, their slow reactivity toward OH radicals, and the difficulties encountered in the IR spectra evaluation, the plots depicted in Figs. 1 to 4 show good linearity. Total conversion of nitrocatechols during the irradiation period was evaluated to range from about 25 % to 55 %. Up to 60 % of the decays of the nitrocatechols and the reference compounds have been attributed to the reaction with OH radicals during the kinetic experiments.

Table 1 presents the kinetic data results from the present study. Relative ratios of k_1/k_2 determined from the kinetic plots (see Figs. 1 to 4), gas-phase rate coefficient (k_1) calculated for the OH reactions with the investigated nitrocatechols, and the average kinetic coefficients (k_1 average) are also indicated in Table 1. Uncertainties for the k_1/k_2 represent 2σ from the linear regression analysis. The k_1 rate coefficient errors are given from a combination of 2σ values

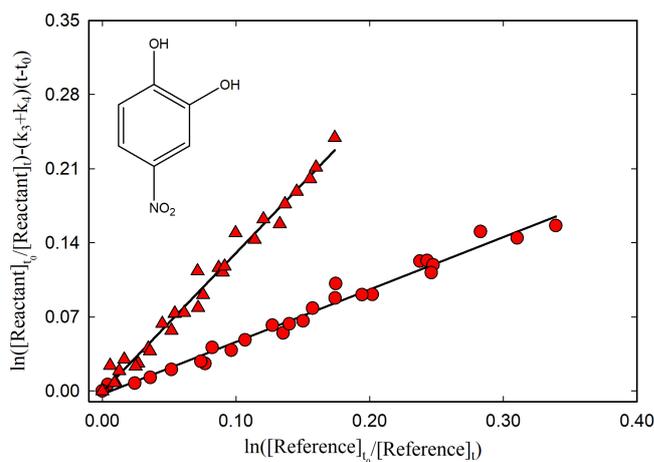


Figure 2. Kinetic plot according to Eq. (1) for the reaction of 4-nitrocatechol with OH radicals relative to dimethyl ether (red circle), using photolysis of $\text{CH}_3\text{ONO}/\text{NO}$ mixture at 365 nm as OH-radical source, and relative to methanol (red triangle), using photolysis of H_2O_2 at 254 nm as OH-radical source, with $(k_3 + k_4) = (1.37 \pm 0.32) \times 10^{-4} \text{ s}^{-1}$.

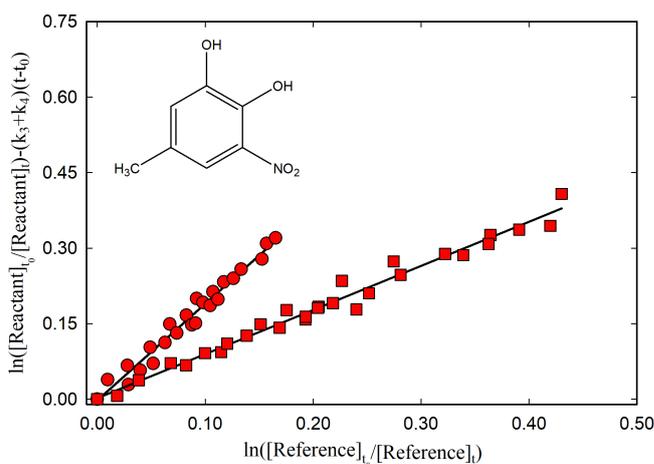
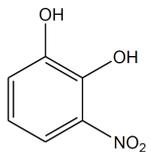
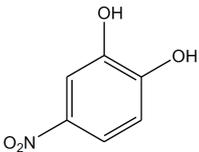
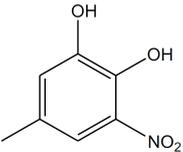
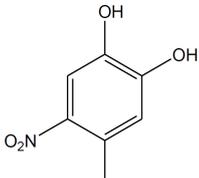


Figure 3. Kinetic plot according to Eq. (1) for the reaction of 5-methyl-3-nitrocatechol with OH radicals relative to dimethyl ether (red circle) and cyclohexane (red square) using photolysis of $\text{CH}_3\text{ONO}/\text{NO}$ mixture at 365 nm as OH radical source, with $(k_3 + k_4) = (4.27 \pm 0.98) \times 10^{-4} \text{ s}^{-1}$.

from the linear regression and an additional error from the reference compound uncertainty. As for the k_1 (average) and its uncertainty, the weighted average was used to estimate the best interval that accommodates the highest confidence of the obtained data. Based on the daytime average concentration of the OH radicals in the atmosphere (Prinn et al., 1995), of $1.6 \times 10^6 \text{ cm}^{-3}$, Table 1 lists also the atmospheric residence time of the studied nitrocatechols. No significant differences were observed for 4-nitrocatechol and 4-methyl-5-nitrocatechol rate coefficient values when NO_x or NO_x -free conditions have been employed.

Table 1. Rate coefficient values of OH radical reactions with all investigated nitrocatechols in the present study.

Compound	Reference	k_1/k_2	k_1 ($10^{-12} \text{ cm}^3 \text{ s}^{-1}$)	k_1 (average) ($10^{-12} \text{ cm}^3 \text{ s}^{-1}$)	τ^* (h)
 3-nitrocatechol	DME	1.17 ± 0.09	3.27 ± 0.69	3.41 ± 0.37	52
	cyclohexane	0.54 ± 0.06	3.46 ± 0.44		
 4-nitrocatechol	DME	0.49 ± 0.03	1.39 ± 0.29	1.27 ± 0.19	137
	methanol	1.31 ± 0.06	1.18 ± 0.24		
 5-methyl-3-nitrocatechol	DME	1.92 ± 0.13	5.37 ± 1.14	5.55 ± 0.48	31
	cyclohexane	0.88 ± 0.05	5.58 ± 0.49		
 4-methyl-5-nitrocatechol	DME	0.29 ± 0.03	0.82 ± 0.18	0.92 ± 0.14	190
	methanol	1.18 ± 0.06	1.06 ± 0.22		

* Calculated by applying $\tau = 1/(k_1 \times [\text{OH}])$ and considering the average daytime OH-radicals concentration ($[\text{OH}]$) of $1.6 \times 10^6 \text{ cm}^{-3}$ (Prinn et al., 1995).

Direct photolysis of the hydroxylated nitroaromatics under similar conditions to those used in this study has been shown to produce HONO (Bejan et al., 2006), similar to the carboxylic acid formation mechanism in the aqueous phase proposed by Alif et al. (1991). Photolysis of HONO is a well-known source of OH radicals. To properly evaluate the photolysis rate coefficients at 365 and 254 nm for the investigated nitrocatechols by accounting for possible interference of the secondary chemistry owing to the OH radicals' reactions, a series of individual experiments were performed using carbon monoxide (CO) as OH scavenger. To override the low reactivity of the scavenger ($k_{\text{CO}+\text{OH}} = 2.41 \times 10^{-13} \text{ cm}^3 \text{ s}^{-1}$) (DeMore et al., 1997), a high amount of CO was introduced into the reactor vessel (of about $1.17 \times 10^{17} \text{ cm}^{-3}$) to ensure a scavenging efficiency of

more than 98 %. Photolysis rates for all four nitrocatechols at 365 nm (J_{365}) were evaluated. At this wavelength, no significant photolysis for 4-nitrocatechol and 4-methyl-5-nitrocatechol was observed, whereas for 3-nitrocatechol and 5-methyl-3-nitrocatechol photolysis frequencies of $(1.53 \pm 0.08) \times 10^{-4}$ and $(1.07 \pm 0.09) \times 10^{-4} \text{ s}^{-1}$, respectively, have been measured (see Fig. S1 in the Supplement). Additionally, because for 4-nitrocatechol and 4-methyl-5-nitrocatechol the reaction rate coefficients were measured under NO_x free condition by using H_2O_2 photolysis as OH-radical precursor, photolysis rate coefficients at 254 nm of these two compounds were evaluated. For 4-nitrocatechol the photolysis rate (J_{254}) was $(6.71 \pm 0.99) \times 10^{-5} \text{ s}^{-1}$ and that for 4-methyl-5-nitrocatechol was $(3.18 \pm 0.32) \times 10^{-5} \text{ s}^{-1}$ (see Fig. S2). However, for 3-nitrocatechol and 5-methyl-3-

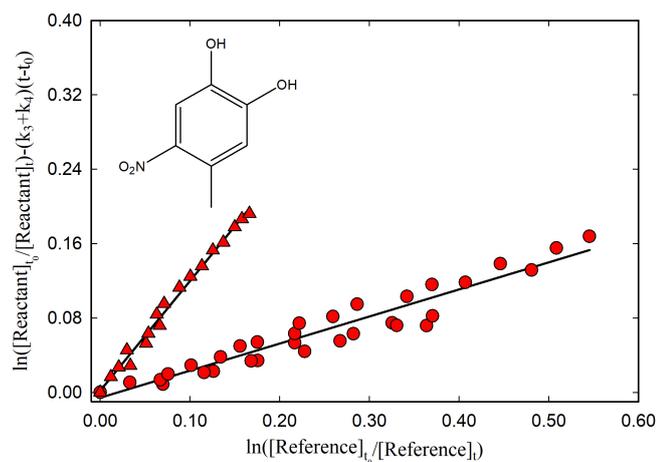


Figure 4. Kinetic plot according to Eq. (1) for the reaction of 4-methyl-5-nitrocatechol with OH radicals relative to dimethyl ether (red circle), using photolysis of $\text{CH}_3\text{ONO}/\text{NO}$ mixture at 365 nm as OH-radical source, and relative to methanol (red triangle), using photolysis of H_2O_2 at 254 nm as OH-radical source, with $(k_3 + k_4) = (1.02 \pm 0.30) \times 10^{-4} \text{ s}^{-1}$.

nitrocatechol corrections have been also performed because of their high wall loss and their photolysis at 365 nm. Photolysis rate coefficients are presented in Table 2, in addition to the average lifetime with respect to photolysis under atmospheric conditions. Photolysis rate coefficients at 365 nm for 3-nitrophenols have been scaled to the atmospheric conditions, as Klotz et al. (1997) have described for 40° N latitude noontime and clear sky conditions, with a factor of 2 obtained from the NO_2 photolysis frequency measured under atmospheric conditions versus the ESC-Q-UAIC reactor ($J_{\text{NO}_2\text{-atmosphere}}/J_{\text{NO}_2\text{-ESC-Q-UAIC chamber}} = 8.5 \times 10^{-3} \text{ s}^{-1}/4.3 \times 10^{-3} \text{ s}^{-1} \approx 2$). The wall loss (k_4) measured for 3NCAT and 5M3NCAT was about $(3.2 \pm 0.4) \times 10^{-4} \text{ s}^{-1}$, whereas for 4NCAT and 4M5NCAT it was $(0.7 \pm 0.3) \times 10^{-4} \text{ s}^{-1}$. However, the wall loss was relatively constant over the different experiments employed in this study. Average quantum yields calculated for 3NCAT and 5M3NCAT over the 350- to 400-nm range were obtained using the theoretical approach described by Hofzumahaus et al. (1999), and the evaluated cross-sections of these two compounds in aqueous solutions are presented in Table S1 in the Supplement. The average quantum yields, calculated for the investigated range ($\varphi_{350-400}$), are: $(7.7 \pm 0.7) \times 10^{-3}$ for 3NCAT and $(4.9 \pm 0.5) \times 10^{-3}$ for 5M3NCAT. The overall errors include contributions from the solution preparation, UV-Vis spectra recording (see Fig. S3a for 3NCAT and Fig. S3b for 5M3NCAT) and the photolysis values. A detailed description of the technique, data collection, evaluation and interpretation is presented in the Supplement. However, preliminary analysis of the residual spectra obtained after irradiation of 3NCAT and 5M3NCAT, presented in Fig. S4, reveals one signal at 2139 cm^{-1} for 3NCAT and

2135 cm^{-1} for 5M3NCAT that could be representative of a ketene-type product.

5 Discussion

The rate coefficients obtained from the kinetic investigations of a series of nitrocatechols with OH radicals are reported for the first time in this study. No significant differences have been observed between the two reference compounds that have been employed to achieve the quality of the rate coefficient values. Based on collision theory and the standard gas kinetic theory, Sørensen et al. (2002) showed that the presence of inorganic aerosols does not influence the value of the OH radicals rate coefficients when relative kinetic technique is used. The OH radical deactivation from collision with surfaces, both with aerosols and chamber, were found to be negligible for all nitrocatechols. However, even if such a process occurs at a higher rate, using the relative kinetic technique, where the k_1/k_2 ratio is experimentally determined independent of the OH radical concentration, no systematic errors arising from the OH-radical concentration would influence the kinetic data.

A comparison with literature data is not possible. However, there are reported rate coefficient values for nitrophenol and methylated nitrophenols reaction with OH radicals (Atkinson et al., 1992; Bejan et al., 2007). Additionally, kinetic rate coefficients of the OH-radical-initiated oxidation of nitronaphthalene were studied by Atkinson et al. (1989). For the OH-radical reactions with the nitroaromatic compounds, however, all the reported values of the rate coefficients in the literature have the same order of magnitude of $10^{-12} \text{ cm}^3 \text{ s}^{-1}$, as those presented in this work.

5.1 Electromeric effect of aromatic ring substituents

Prior to initiating discussions of the kinetic results, we should emphasize the influence of the substituent position on the aromatic ring for all the nitrocatechols and their effect on the reactivity. The electromeric effects influencing reactivity could be easily noticed on the structure of infrared spectra. Figure 5a presents the gas-phase infrared spectra of catechol (CAT), 3NCAT, and 4NCAT, and Fig. 5b the gas-phase infrared spectra of 4-methylcatechol (4MCAT), 5M3NCAT, and 4M5NCAT. Infrared spectra clearly show the presence of the H-bonds between OH (vibrations being marked with * and **) and NO_2 group substituents attached to the aromatic ring in the nitrocatechols as follows: the compounds 3NCAT and 5M3NCAT exhibit an H-bond between the OH and one of the O atoms from the NO_2 vicinal group; 4NCAT presents an H-bond between the H from the C1-linked OH and the O atom from the C2-linked OH; and 4M5NCAT reveals an H-bond between the H from the C2-linked OH and the O atom from the C1-linked OH in 4M5NCAT similar to 4NCAT. The existence of similar structure of vicinal OH groups could be observed in catechol and 4-methylcatechol.

Table 2. Rate coefficient values for the photolysis of all nitrocatechols investigated in this study scaled to atmospheric conditions and their average atmospheric lifetimes owing to photolysis (at 365 nm).

Compound	$J_{(\text{NCAT})} - 254 \text{ nm}$ (10^{-4} s^{-1})	$J_{(\text{NCAT})} - 365 \text{ nm}$ (10^{-4} s^{-1})	$\tau_{h\nu}^*$ (at 365 nm) (h)
3NCAT	not measured	3.06 ± 0.16	~ 0.9
4NCAT	0.67 ± 0.01	< 0.1	> 27.8
5M3NCAT	not measured	2.14 ± 0.18	~ 1.3
4M5NCAT	0.32 ± 0.03	< 0.1	> 27.8

* Calculated by applying $\tau = 1/J_{(\text{NCAT})}$, where $J_{(\text{NCAT})}$ at 365 nm has been estimated from the experimental determined $J_{(\text{NCAT})}$ (ESC-Q-UAIC chamber) and using a scaled factor derived from the NO_2 photolysis frequency for the atmospheric conditions described by Klotz et al. (1997) (see in text).

However, these H-bonds are not present in CAT and 4MCAT, as both OH groups are highlighted in the gas-phase IR spectra. H-bonds occurring in the nitrocatechols should be evaluated in light of the existing nitro group, which acts with its deactivating electromeric $E-$ effect on the ring, withdrawing the electrons from its *-ortho* and *-para* positions. For the 3-nitrocatechols, the chemical structures reveal the presence of one OH group placed between the OH and NO_2 groups. This hydroxyl group will be affected by the electron withdrawing effect of NO_2 , and as a consequence, the weakened O-H bond results in an H-atom that is more susceptible to forming an H-bond with the O from the NO_2 group. This bond can also be observed in the IR spectra of 2-nitrophenols (Fig. S5) (Olariu et al., 2002, 2013; Bejan, 2006; Bejan et al., 2007) and is considered to be responsible for SOA and HONO formation via H-atom transfer during gas-phase photolytic processes (Bejan et al., 2006, 2020). As shown in the spectra from Fig. 5a and b, the structures of CAT and 4MCAT exhibit the presence of OH groups one next to the other and no visible H-bond between these substituents. This suggests that for 1,2-dihydroxybenzenes the $E+$ effect is less diminished through the presence of the vicinal OH groups.

In Fig. 6, 3NCAT has been chosen as a master example to emphasize the electromeric effects manifested by NO_2 and OH groups on the aromatic ring, in order to also identify their possible implications for molecule stability. For the 3NCAT structure, one could observe that the OH substituent next to the NO_2 group would produce an H bond that is shielded to manifest its $E+$ effect on the ring. As reported by Bejan et al. (2006, 2020), photolysis of methylated 2-nitrophenols leads to HONO formation via H-transfer from the hydroxyl group to its vicinal nitro group. For 4NCAT and 4M5NCAT, where both OH are linked in the H-bonds, no significant photolysis at 365 nm and no $E+$ effect to enhance electron density on the aromatic ring occur.

5.2 Reactivity trends

It is well known that the presence of a nitro group has a high deactivating effect on the aromatic ring as the reactivities of nitroaromatic compounds are approximately 5–20 times

lower than their parent compounds. Some OH rate coefficients at 298 K for AHs and their corresponding NAHs are presented in Table 3. The reactivity of catechols against OH radicals drops even more than two orders of magnitude when a nitro group is attached to the ring (Olariu et al., 2000). Such a sudden loss in reactivity is due to the inhibition of the addition channel in aromatic hydrocarbons as discussed in the present study. Ring-retaining product studies from the gas-phase, OH-radical-initiated oxidation of phenol and cresols show that the addition pathway is the dominant oxidation mechanism, occurring in between 65 % and 90 % of the overall reactions, nitrophenols being formed via phenolic H-atom abstraction channel accounting for 5 % to 10 % (Olariu et al., 2002). Gas-phase reactions with NO_3 radicals lead to a yield of ~ 50 % of nitrocatechols and almost 100 % of nitric acid, suggesting that the H-atom abstraction channel represents at least half of the overall reactions (Olariu et al., 2013). In the gas-phase oxidation of catechol with OH radicals under high NO_x conditions, Finewax et al. (2018) found 4NCAT in the SOA composition with a molar yield of 30 %. At least 90 % molar yield was obtained in their study in the oxidation initiated by NO_3 radicals. These results suggest that the H-atom abstraction channel might make a greater contribution to catechol reactivity rather than the abstraction channel from phenol or cresol oxidation.

The gas-phase rate coefficients in the investigated series of nitrocatechols vary as follows: $k_{5\text{M3NCAT}} > k_{3\text{NCAT}} > k_{4\text{NCAT}} \sim k_{4\text{M5NCAT}}$. 4NCAT and 4M5NCAT could assume a small effect from the methyl group but shows little distinction from experimental determined rate coefficient values from this study. A distribution of possible free attack sites by OH radicals on the aromatic ring, accounted for also by the substituent's effects, is represented in Fig. 7a. The NO_2 group deactivates the aromatic ring toward OH attack by the $E-$ effect. By its $I+$ effect, the $-\text{CH}_3$ substituent will probably influence the electrophilic addition, although its effect is diminished by the $E-$ from the $-\text{NO}_2$ substituent. According to Fig. 7a, the most stable adduct is the 4M5NCAT-OH, and one could propose an OH radical addition to the third position of the ring. However, this is

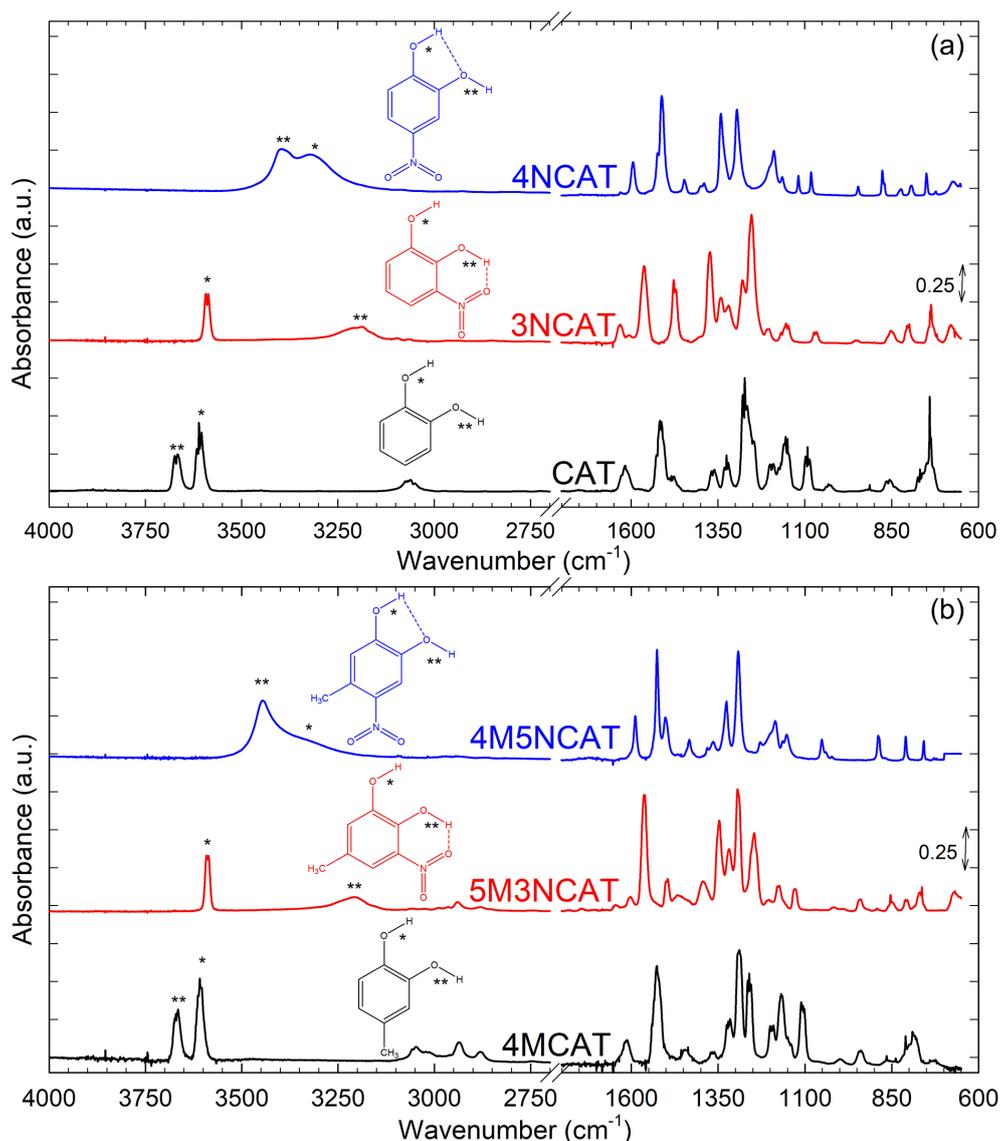


Figure 5. (a) Gas-phase IR spectra of catechol (CAT), 3NCAT, 4NCAT; (b) Gas-phase IR spectra of 4-methylcatechol (4MCAT), 5M3NCAT, and 4M5NCAT. Evidence for the intramolecular H-bond occurrence between the substituents marked with */** : catechol and 4MCAT – both free OH present in the spectra; 3NCAT and 5M3NCAT – only one free OH; 4NCAT and 4M5NCAT – both OH involved in the H-bond.

not the case as the experimental data show that 4M5NCAT is the least reactive compound in this series, because of the substituent's interactions (see Fig. 7b). Differences between the reactivity of 3NCAT and 5M3NCAT in reaction with an electrophile suggest that two determining pathways might occur: addition to the ring in the NO_2 -deactivated fourth or sixth position oriented by the $E+$ electromeric effect of OH (C1) and the $I+$ effect from CH_3 , and H-atom abstraction from the C1-linked OH group. Although the difference between 3NCAT and 5M3NCAT is comparable with the difference between 2-nitrophenol (2NPh) and 4-methyl-2-nitrophenol (4M2NPh), we may consider that the $E+$ effect of OH (C1) is less present owing to the same orien-

tation of the $E-$ effect of NO_2 . Comparing the reactivity of 2NPh with those of methyl-2-nitrophenols, an increase in the reactivity of 2 to 7 times is observed, which is attributed to CH_3 group presence and orientation. The most reactive compound is 5-methyl-2-nitrophenol (5M2NPh) with a $k_{5\text{M2NPh}} = 6.72 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, where the methyl $I+$ effect is manifested on both *-ortho* positions relative to itself and is not affected by the NO_2 $E-$ deactivating effect. In comparison, 4M2NPh, $k_{4\text{M2NPh}} = 3.59 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, both *-ortho* positions to methyl are likely deactivated by the NO_2 . In the methylated nitrophenol series, Bejan et al. (2007) found that the least reactive is 6-methyl-2-nitrophenol (6M2NPh), $k_{6\text{M2NPh}} = 2.70 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, with just one

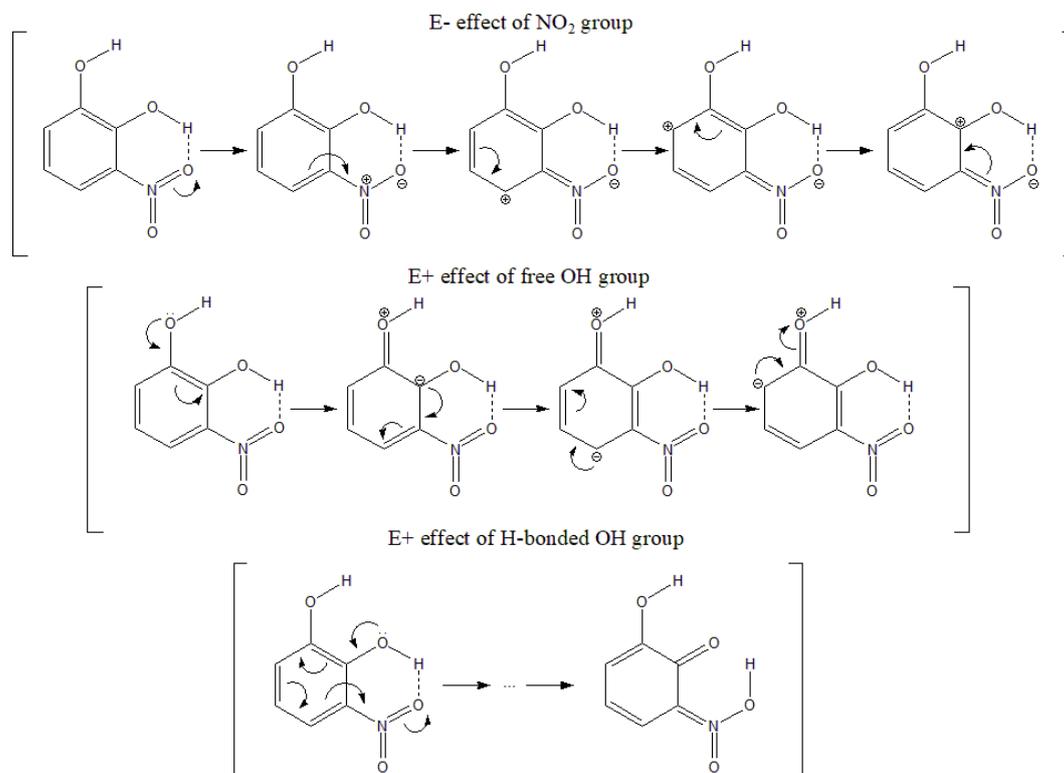


Figure 6. Electromeric effects of NO₂ and OH substituents from the 3NCAT molecule.

Table 3. Rate coefficients at 298 K for aromatic hydrocarbons (AHs) and their correspondent nitroaromatic (NAHs) derivatives for the gas-phase reaction with the OH radical.

Aromatic hydrocarbon (AHs)	k_{OH} ($10^{-12} \text{ cm}^3 \text{ s}^{-1}$)	Nitroaromatic hydrocarbon (NAHs)	k_{OH} ($10^{-12} \text{ cm}^3 \text{ s}^{-1}$)	$k_{\text{NAHs}}/k_{\text{AHs}}$ (%)
Benzene	1.22 ^a	nitrobenzene	0.14 ^{a,b}	11.48
Toluene	5.63 ^a	<i>m</i> -nitrotoluene	1.20 ^a	21.31
Phenol	26.3 ^{a,d}	2-nitrophenol	0.90 ^c	3.42
<i>o</i> -Cresol	41 ^{a,d}	6-methyl-2-nitrophenol	2.70 ^e	6.59
<i>m</i> -Cresol	68 ^{a,d}	3-methyl-2-nitrophenol	3.69 ^e	5.43
		5-methyl-2-nitrophenol	6.72 ^e	9.88
<i>p</i> -Cresol	50 ^{a,c}	4-methyl-2-nitrophenol	3.59 ^e	7.18
Catechol	104 ^f	3-nitrocatechol	3.41 ^g	3.28
		4-nitrocatechol	1.27 ^g	1.22
4-Methylcatechol	156 ^f	5-methyl-3-nitrocatechol	5.55 ^g	3.56
		4-methyl-5-nitrocatechol	0.92 ^g	0.59
Naphthalene	23.00 ^a	1-nitronaphthalene	5.40 ^h	23.48
		2-nitronaphthalene	5.60 ^h	24.35

^a Calvert et al. (2002). ^b Witte et al. (1986). ^c Atkinson et al. (1992). ^d Atkinson (1989). ^e Bejan et al. (2007). ^f Olariu et al. (2000). ^g This work. ^h Atkinson et al. (1989).

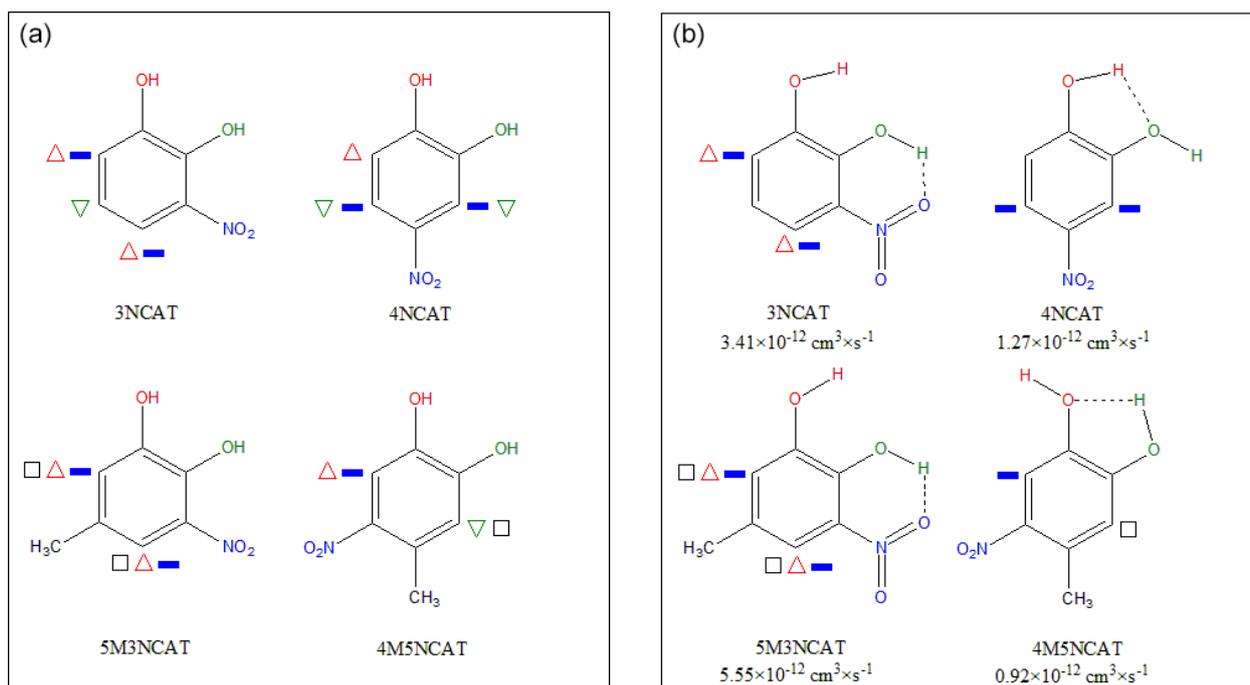


Figure 7. Electromeric effect distribution on the aromatic ring for the all investigating nitrocatechols in the present study: (a) active sites toward electrophilic attack by hydroxyl (1st OH: red triangle/2nd OH: green triangle), nitro (blue bar), and methyl (black square) groups; (b) internal interactions of the substituents and consequences on the reactivity based on the gas phase FT-IR spectra.

-ortho and one *-para* position that are both deactivated by NO_2 . Following the trend for 3-nitrocatechols with respect to the internal electromeric effects of the OH and NO_2 groups and considering the reactivity trend in methyl-2-nitrophenols (Bejan et al., 2007) and the implication of the methyl group (Klotz et al., 1998; Bejan et al., 2012; Aschmann et al., 2013) on the addition of OH radicals to the (hydroxylated) aromatic hydrocarbons (Olariu et al., 2000), the rate coefficient for 6M3NCAT (not measured in this study) was estimated to fall between $(5.0 \text{ and } 6.5) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$, similar to 5M3NCAT.

For 4NCAT and 4M5NCAT, the difference in the reactivity is not obvious as both OH groups involved in the H-bond have their $E+$ effect shielded by the H-bond. If the OH group could manifest its electromeric effect, the most reactive compound should be 4M5NCAT. However, the supplementary $I+$ effect of methyl group, *-para* oriented from the C1-linked OH in 4M5NCAT, seems to strengthen the C1 O-H bond, making the H abstraction harder. Even if H-abstraction from CH_3 could be an option for 4M5NCAT, the overall reactivity of this compound is smaller or at least the same as that of 4NCAT, H-abstraction from the OH group C1-linked being the main degradation path. Seeking the trend in 4-nitrocatechols and 5-nitrocatechols based on these experimental observations, we could estimate that for 3M4NCAT the OH rate coefficient falls between $(0.5 \text{ and } 0.9) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ and for 3M5NCAT between $(1.0 \text{ and } 1.4) \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$. However, for all 4-nitrocatechols and

5-nitrocatechols, the rate coefficients could be considered to be around $1 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$. These estimations should be treated carefully, as few experimental data are available at this time.

5.3 Comparison with SAR estimated values

There are two Structure-Activity-Relationship (SAR) models for estimating gas-phase OH-kinetic rate coefficients of aromatic hydrocarbons, both based on substituent factor analysis, that influence the reactivity under atmospheric relevant conditions. The general model proposed by Kwok and Atkinson (1995) (used mainly for aliphatic hydrocarbons) considers four cumulative reaction pathways for the OH-initiated oxidations: H-atom abstraction from ring substituents, OH-radical addition to the unsaturated aliphatic (double or triple) bond of the substituents, OH-radical addition to the aromatic ring and OH-radical interactions with ring heteroatoms. The rate coefficient for the total reaction is assumed as a sum of rate coefficients from these four pathways. As for aromatics, this model proposes standard values for H-atom abstraction for the group rate coefficients and substituent factors. For addition to the aromatic ring, the Hammett constants for electrophilic addition, σ^+ , given by Brown and Okamoto (1958), were proposed by Zetzsch (1982) and updated by Kwok and Atkinson (1995) to fit to the equation $\log(k_{\text{add}}/\text{cm}^3 \text{ s}^{-1}) = -11.71 - 1.34 \times \sum \sigma_i^+$, where the most negative value of the sum $\sum \sigma_i^+$

has to be considered alone and any interaction among the substituents is neglected. The SAR estimation model from Jenkin et al. (2018a, b) updates the factors used in the calculation of rate coefficients and branching ratios for the gas-phase reactions of the OH radicals with monocyclic aromatic compounds. A scaling parameter for the OH-radical addition to each carbon atom in the aromatic cycle relative to positions of the ring substituents was introduced. Additionally, for the addition pathway, 17 possible ring substituent adjustment factors were introduced, including nitro and hydroxyl groups. In the H-atom abstraction pathway from aliphatic substituents, another scaling factor for group rate coefficients of $\exp(140/T)$ is recommended to be used when alkyl substituents are present in *-ortho* or *-para* position relative to the attack site. Based on previous gas-phase kinetic and product studies on OH-radical-initiated oxidation of hydroxylated aromatic compounds, employed for cresols (Olariu et al., 2002; Coeur-Tourneur et al., 2006), methoxyphenols (Coeur-Tourneur et al., 2010b, a), catechols (Olariu et al., 2000; Finewax et al., 2018), and nitrophenols (Bejan et al., 2007, 2020), an abstraction rate coefficient from the phenolic OH of $k_{\text{abs}}(\text{Ph-OH}) = 2.6 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ was proposed by Jenkin et al. (2018b), a rate coefficient that is 20 times larger than those observed in aliphatic alcohols (Jenkin et al., 2018a). This has a great impact on the overall reactivity and the branching ratios of aromatic hydroxylated monocyclic compounds. Observations from the kinetic data of nitro-containing aromatic hydrocarbons result in re-evaluation of the H-atom abstraction channel from the OH substituent, and a more suitable value for the reactions occurring at 298 K is $k_{\text{abs}}(\text{Ph-OH}) = 1.4 \times 10^{-13} \text{ cm}^3 \text{ s}^{-1}$. EPI Suite – AOPWIN software 4.11 was developed based on Kwok and Atkinson's (1995) observations by the United States Environmental Protection Agency. Data in Table 4 show that each SAR tends to overestimate the OH gas-phase reactivity of hydroxylated nitrobenzenes. Along with nitrocatechols rate coefficient values obtained in the present study, Table 4 also lists the experimental kinetic rate coefficients for 2-nitrophenols determined by Bejan et al. (2007) and the estimated SAR values. The 2-nitrophenols estimated rate coefficient values of Jenkin et al. (2018b) fit better with the experimental data than other SARs, owing to the use of substituent adjustment factors $R(\Phi)$ for the substituent factors $F(\Phi)$ updated to the available literature and the rescaled abstraction rate coefficient previously assigned to -OH groups in aliphatic compounds of $k_{\text{abs}}(\text{Ph-OH}) = 1.28 \times 10^{-12} \times \exp(-660/T) \text{ cm}^3 \text{ s}^{-1}$ ($1.4 \times 10^{-13} \text{ cm}^3 \text{ s}^{-1}$ at 298 K). In the case of nitrocatechols, Jenkin et al. (2018b) SAR values are a factor of 1.5 higher than the values of Kwok and Atkinson (1995). The difference between the estimated values is considered to appear because of the k_{add} estimation in Kwok and Atkinson's (1995) SAR, where the electrophilic substituent coefficient σ^+ for the OH oriented in *-meta* position has no attributed contribution. For 1,2-dihydroxybenzenes molecules, the addition

pathway might count only for one phenolic hydroxyl. This is in agreement with the experimental explanation given for 3-nitrocatechol's reactivity, where the $E+$ effect of the OH group from C2 has no electromeric effect on the aromatic ring as is visible in infrared spectra. Consistency between the rate coefficient values is given as both SARs assume that the addition pathway is at least 94 % of the overall reactivity. Based on the observations of Bejan et al. (2007) on methylated 2-nitrophenols, this assumption is reasonable.

As the data in Table 4 show, no differences were observed between the rate coefficient values estimated by the EPI Suite – AOPWIN and Kwok and Atkinson's (1995) approach as the programme is based on the model calculation proposed study. However, for nitrocatechols there is a difference between the rate coefficients values. This difference is observed because the programme could not choose the correct parameters of the nitro group to properly calculate $\Sigma\sigma^+$. Instead of choosing the lowest value for NO_2 substituent *-meta* oriented, it chooses the electrophilic substituent group constant from *-ortho* and *-para*, that gives the highest $\Sigma\sigma^+$, not the lowest, as is recommended. However, in further discussions we will no longer consider EPI Suite – AOPWIN software data.

In SAR approaches, the estimated rate coefficients values do not consider the interactions between the nitro and its vicinal OH group or between OH groups, as they all have a tendency to overestimate the gas-phase OH rate coefficients for hydroxylated nitro monocyclic aromatics. However, Jenkin et al. (2018b) proposed a substituent adjustment factor for NO_2 groups of 0.024 for the *-ortho* and *-para* relative to each of the addition sites on the ring and of 0.07 for the *-ipso* and *-meta* positions. In Kwok and Atkinson (1995), the SAR estimations assign the NO_2 group substituent factor of Brown and Okamoto (1958) of +0.790 in *-ortho* or *-para* and of +0.674 for *-meta* or *-ipso*. All of these factors are calculated based on previous experimental data to fit alkyl substituted or low substituted aromatic compounds. Although the updated SAR of Jenkin et al. (2018b) is more complex, interactions between the substituents are not taken into consideration. These limitations lead in the case of 4NCAT or 4M5NCAT to different ratios of 5 and 18 times higher. For the 3NCAT and 5M3NCAT the estimated rate coefficient values are more consistent with the present experimental data, values being only ~ 2 times larger. Differences between experimental values of 4NCAT and 4M5NCAT in comparison with those estimated by SAR models suggest that the addition pathway is of less importance than the H-atom abstraction from hydroxyl groups owing to the inhibition of the $E+$ effect of phenolic hydroxyls, and limited contribution from the presence of the CH_3 group.

If substituent interactions are taken into account by: (1) excluding the OH substituent effect on the addition channel (one substituent effect in 3NCAT, 5M3NCAT, and both substituents in 4NCAT and 4M5NCAT); (2) removing one H-atom abstraction channel in all the nitrocate-

Table 4. Experimental and SAR-estimated rate coefficients of monoaromatic hydroxylated nitro compounds with the radicals. In brackets are the ratios of the estimated values versus experimental data ($k_{\text{calc}}/k_{\text{exp}}$).

Compound	Experimental data	SAR				
		EPI Suite – AOPWIN	Kwok and Atkinson (1995)	Jenkin et al. (2018b) $k_{\text{abs}}(\text{Ph-OH})$		this work
				2.6×10^{-12}	0.14×10^{-12}	
		$k_{\text{(OH)}} \times 10^{12} \text{ (cm}^3 \text{ s}^{-1}\text{)}$				
3NCAT	3.41 ± 0.37	3.15 (0.92)	4.45 (1.30)	11.83 (3.47)	6.91 (2.03)	3.95 (1.16)
4NCAT	1.27 ± 0.19	3.15 (2.48)	4.45 (3.50)	11.83 (9.31)	6.91 (5.44)	2.36 (1.86)
5M3NCAT	5.55 ± 0.48	5.65 (1.02)	8.02 (1.45)	18.21 (3.28)	11.81 (2.13)	6.15 (1.11)
4M5NCAT	0.92 ± 0.14	7.90 (8.59)	11.29 (12.27)	22.99 (25.00)	16.60 (18.04)	3.71 (4.03)
3M2NPh	$3.69 \pm 0.70^*$	11.15 (3.02)	11.15 (3.02)	7.19 (1.95)	4.73 (1.28)	2.00 (0.54)
4M2NPh	$3.59 \pm 1.17^*$	5.38 (1.50)	5.38 (1.50)	7.17 (2.00)	3.24 (0.90)	1.41 (0.39)
5M2NPh	$6.72 \pm 2.14^*$	11.15 (1.66)	11.15 (1.66)	7.19 (1.07)	4.73 (0.70)	2.00 (0.30)
6M2NPh	$2.70 \pm 0.57^*$	5.38 (1.99)	5.38 (1.99)	7.17 (2.66)	3.24 (1.20)	1.41 (0.52)

* Values from Bejan et al. (2007): 3M2NPh- 3-methyl-2-nitrophenol; 4M2NPh- 4-methyl-2-nitrophenol; 5M2NPh- 5-methyl-2-nitrophenol; 6M2NPh- 6-methyl-2-nitrophenol.

echols; (3) considering the methyl group inhibiting the H-atom abstraction channel in 4M5NCAT; and (4) using a new neighbouring group factor for α -H-atom abstraction from the deactivated phenolic group $F(-\text{Ph deactivated})=0.5$, then the automated mechanism construction proposed by Jenkin et al. (2018b) would emphasize new estimated rate coefficient values approaching the experimental data obtained in this study as follows: for 3NCAT a rate of $3.95 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ with a $k_{\text{add}}/k_{\text{tot}}=0.67$, for 5M3NCAT a rate of $6.15 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ with a $k_{\text{add}}/k_{\text{tot}}=0.73$, for 4NCAT a rate of $2.36 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ with a $k_{\text{add}}/k_{\text{tot}}=0.45$ and for 4M5NCAT a rate of $3.71 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ with a $k_{\text{add}}/k_{\text{tot}}=0.69$. However, based on the study of Jenkin et al. (2018b) and the behaviour of the investigated nitrocatechols it is probably better to use a value of 0.5 as a new parameter for the H-atom abstraction for the deactivated phenyl group $F(-\text{Ph deactivated})$ when $k_{\text{abs}}(\text{Ph-OH})=2.6 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ is considered. In this case, the H-atom abstraction from phenolic hydroxyl will gain significant importance and will be in agreement with previous kinetic studies of hydroxylated aromatic compounds. The new estimated values obtained within the present work are listed in Table 4 besides those of methylated nitrophenols, for which systematic lower values are reported. Table S2 in the Supplement presents the revised estimated k_{OH} values for several nitroaromatic compounds using the Kwok and Atkinson (1995) SAR model updated with recent observations. Accounting for these experimental observations, revisions for the SAR model proposed by Jenkin et al. (2018b) were also performed. Figure 8 presents the correlation analyses between the experimental, the original estimated, and the revised estimated rate coefficient values. Thus, Fig. 8a presents such a correlation versus the variation of the $\Sigma\sigma^+$ calculated without substituents interactions,

whereas Fig. 8b presents this correlation versus the variation of the $\Sigma\sigma^+$ calculated when substituent interactions are taken into account. Data presented in both figures show that the main data cluster would be shifted to positive $\Sigma\sigma^+$ when substituent interactions are considered in compliance with experimental values.

Figure 8c reflects the correlations between the estimated and experimental determined rate coefficient values for the nitrophenols and nitrocatechols. From Fig. 8c it is evident that the estimated revised values show reduced spread throughout the correlation lines of interest generated using the revised Jenkin et al. (2018b) SAR. Moreover, the correlation between estimated k_{OH} values derived from using group factors Jenkin et al., (2018b) versus estimated k_{OH} values derived using electrophilic substituents (Kwok and Atkinson, 1995) is presented in Fig. 8d. From Fig. 8d it is obvious that the original estimate values from both SAR methods present a poor correlation compared with the case when the same estimated k_{OH} values were generated after the revision based on the newly obtained experimental data. The corrections introduced to account for the presence and influence of the H-bond on the rate coefficient values are reflected by the improvement in the Pearson coefficient value arising from $r=0.321$ for the original estimates to $r=0.843$ for the revised estimates. However, in the revised data the experimental rate constant values for nitrophenols seem to be underestimated and those for nitrocatechols overestimated. Under these circumstances, we believe that new experimental values for the rate coefficients of nitroaromatics with OH radicals would further help to better estimate the rate coefficient values for nitroaromatic compounds.

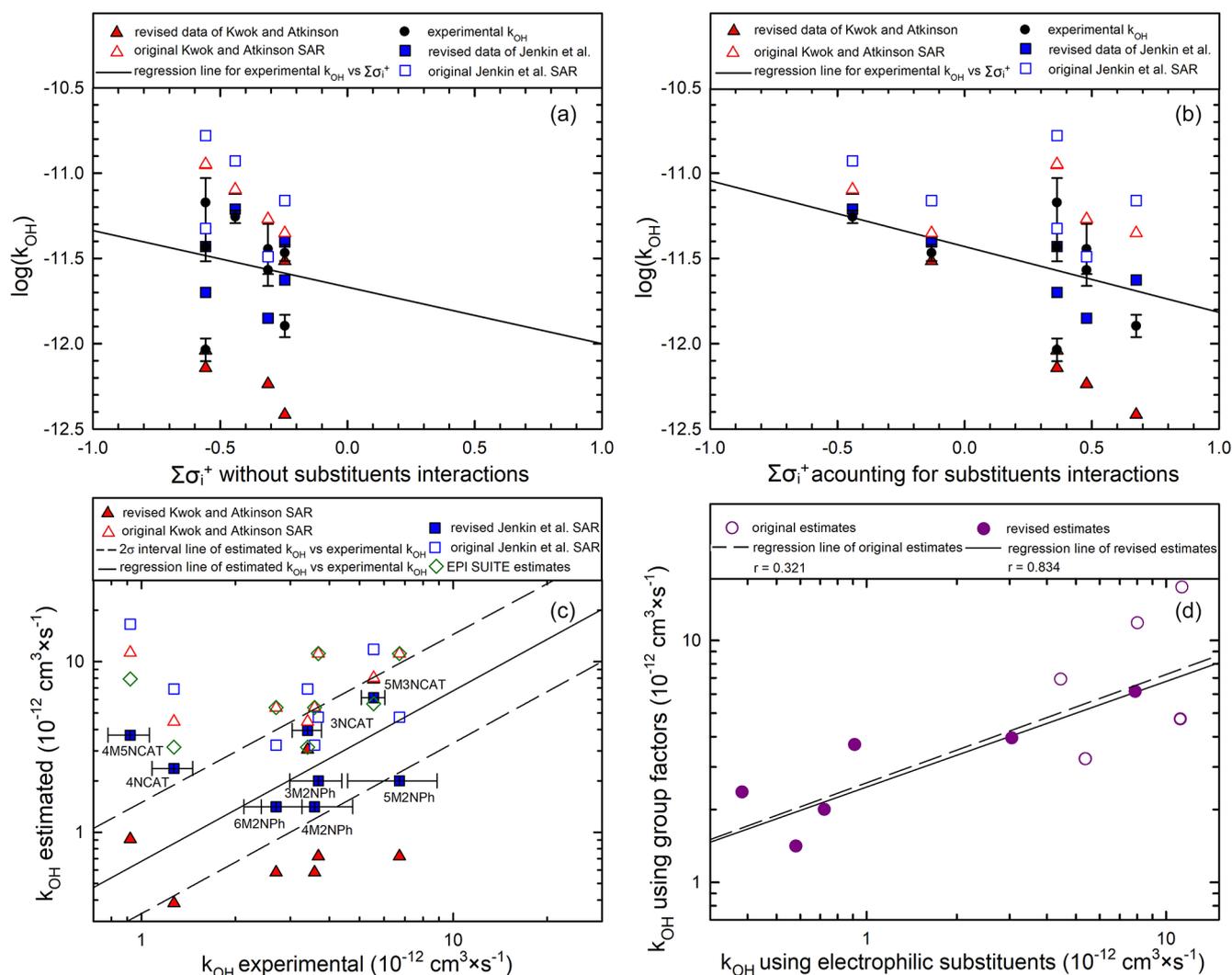


Figure 8. Correlation analyses between the experimental, the original estimated, and the revised estimated OH gas-phase rate coefficients of nitrophenols and nitrocatechols.

5.4 Atmospheric implications

The estimated photolysis frequencies for the 3NCAT and 5M3NCAT extrapolated to atmospheric conditions (see Table 2) are comparable with those proposed by Bejan et al. (2007) for 2-nitrophenols. Photolysis of nitrocatechols with NO_2 and OH groups in vicinal position would be the dominant atmospheric sink, being almost 10 times faster than the reaction with OH radicals. Similar values of photolysis frequencies have been also reported for nitronaphthalenes and methylnitronaphthalenes (Phousongphouang and Arey, 2003). Based on the obtained data, it may be suggested that all the nitroaromatic compounds that have the NO_2 group in the vicinal position with a large electron density area might easily undergo photolysis, which may be the main sink pathway in the atmosphere. Similar to 4-nitro and 5-nitrocatechols, dry or wet deposition seems to be the most

probable atmospheric removal process, with little or no photolysis in the boundary layer and with high chemical residence time regarding atmospheric oxidative degradation initiated by OH radicals.

6 Conclusions

Gas-phase reaction rate coefficients of OH radicals with four nitrocatechols have been investigated for the first time by using ESC-Q-UAIC chamber facilities. The reactivity of all investigated nitrocatechols is strongly influenced by the formation of intramolecular H-bonds that are directly influenced by the deactivating $E-$ effect of the NO_2 group. For the 3-nitrocatechol compounds, the electromeric effect of the “free” OH group is diminished by the deactivating $E-$ effect of NO_2 group. Thus, the rate coefficients of 3-nitrocatechol compounds are higher than the rate coefficient values of

4-nitrocatechol or 4-methyl-5-nitrocatechol, where no electronic effect of the OH group seems to be active. As a consequence, in 4-nitro or 5-nitrocatechols, the H-abstraction pathway seems to play an important role for the entire OH-radical reaction, whereas in the 3-nitrocatechol case the OH-addition path remains dominant.

For 3-nitrocatechols the atmospheric average lifetime seems to be controlled by the photolysis processes, similar to 2-nitrophenols, with an average day-time lifetime of about 1 h. Reaction with OH radicals is the main atmospheric degradation process for 4-nitro and 5-nitrocatechols. Atmospheric day lifetimes of 4-nitro and 5-nitrocatechols are longer than 48 h because of OH-radical-initiated oxidation, being susceptible to long-range transport under atmospheric conditions.

As presented in the current study, the automated mechanism construction for OH rate coefficient SAR estimations is further recommended to consider the internal interactions between vicinal substituents. Moreover, all approaches tend to overestimate the gas-phase reactivity of highly substituted nitro-containing aromatic compounds. Use of an average rate coefficient for the H-atom abstraction from the OH group $k_{\text{abs(Ph-OH)}} = 2.6 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$ is encouraged, along with a scaling parameter for the deactivated aromatic phenyl group, estimated to be 0.5 at 298 K. Also, the substituent adjustment factors $R(\Phi)$, relative to the $F(\Phi)$ values for nitro group should be re-evaluated, $R(\Phi)$ average for all possible addition sites being ~ 0.009 . However, more studies are required to support the proposed reactivity for nitrocatechols and to progress the knowledge of gas-phase aromatic chemistry further.

Data availability. Additional relevant data and supporting information are given in the Supplement. All data used in this study are available upon request from the corresponding author.

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