Significant formation of sulfate aerosols contributed by the heterogeneous drivers of dust surface

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Abstract. The importance of dust heterogeneous oxidation in the removal of atmospheric SO₂ and formation of sulfate aerosols is not adequately understood. In this study, the Fe-, Ti-, and Al-bearing components, Na⁺, Cl⁻, K⁺, and Ca²⁺ of the dust surface, were discovered to be closely associated with the heterogeneous formation of sulfate. Regression models were then developed to make a reliable prediction of the heterogeneous reactivity based on the particle chemical compositions. Further, the recognized gas-phase, aqueous-phase, and heterogeneous oxidation routes were quantitatively assessed and kinetically compared by combining the laboratory work with a modelling study. In the presence of 55 µg m⁻³ airborne dust, heterogeneous oxidation accounts for approximately 28.6 % of the secondary sulfate aerosols during nighttime, while the proportion decreases to 13.1 % in the presence of solar irradiation. On the dust surface, heterogeneous drivers (e.g. transition metal constituents, water-soluble ions) are more efficient than surface-adsorbed oxidants (e.g. H₂O₂, NO₂, O₃) in the conversion of SO₂, particularly during nighttime. Dust heterogeneous oxidation offers an opportunity to explain the missing sulfate source during severe haze pollution events, and its contribution proportion in the complex atmospheric environments could be even higher than the current calculation results. Overall, the dust surface drivers are responsible for the significant formation of sulfate aerosols and have profound impacts on the atmospheric sulfur cycling.

1 Introduction

As an important component of atmospheric particulate matter, sulfate exerts profound impacts on the Earth’s climate system, air quality, and public health (Seinfeld and Pandis, 2016; W. Wang et al., 2021). The rapid formation of sulfate was proven to be largely responsible for London fog and Beijing haze (G. Wang et al., 2016; Cheng et al., 2016). Secondary sulfate aerosols originate predominately from the conversion of sulfur dioxide (SO₂) via the gas-phase oxidation in gaseous environments, aqueous-phase oxidation in liquid media, and heterogeneous oxidation on aerosol surfaces (Ravishankara, 1997; Mauldin III et al., 2012; Su et al., 2020; T. Liu et al., 2021). In recent years, the newly found sulfate formation pathways were kinetically compared with the documented ones to evaluate the relative importance of them (Cheng et al., 2016; Gen et al., 2019; T. Liu et al., 2020; X. Wang et al., 2020). In addition, the reported oxidation channels were compared with each other by aerosol observations or modelling investigations (Berglen et al., 2004; Sarwar et al., 2013; He et al., 2018; Ye et al., 2018; Fan et al., 2020; Tao et al., 2020; Zheng et al., 2020; Song et al., 2021;
Tilgner et al., 2021; Y. Liu et al., 2021; Gao et al., 2022; S. Wang et al., 2022; Ye et al., 2022). In general, these studies emphasized the importance of certain newly discovered aqueous-phase processes or compared the contributions from the documented gas- and aqueous-phase pathways. Nevertheless, heterogeneous reaction was scarcely involved in the discussion, thus hindering the deeper understanding of the atmospheric relevance of aerosol surfaces.

Heterogeneous reaction alters the concentrations of gas-phase SO$_2$ and particle-phase sulfate, and its atmospheric influences were considered by observation and modelling works (Fairlie et al., 2010; Alexander et al., 2012; Chen et al., 2017; K. Wang et al., 2021). As summarized by Table S1, when simulating the sulfate burst events, researchers observed the positive feedbacks after implementing a heterogeneous mechanism in the WRF-Chem (G. Li et al., 2017), GEOS-Chem (Shao et al., 2019), CAMx (Huang et al., 2019), and CMAQ (S. Zhang et al., 2019) models. However, these improved models highlighted the heterogeneous oxidation motivated by the surface-adsorbed oxidants rather than the heterogeneous drivers of aerosol surfaces. To address the knowledge gap, the revised GEOS-Chem (Y. Wang et al., 2014) and WRF-CMAQ (Zheng et al., 2015) models considered the heterogeneous oxidation driven by aerosol surfaces and successfully reproduced the rapid sulfate formation. Xue et al. (2016) moved forward to develop an observation-based model and found that a heterogeneous pathway contributed up to one-third of the secondary sulfate in a typical haze–fog event. However, the “aerosol surface” mentioned in the previous works was not distinguished by its physical or chemical properties (Y. Wang et al., 2014; Zheng et al., 2015; Xue et al., 2016), thereby making it difficult to compare the atmospheric importance of the diverse airborne surfaces. As discussed above, it is of great importance to investigate the heterogeneous drivers of one selected aerosol surface and further evaluate the atmospheric significance of the relevant heterogeneous pathway.

The resurgence of sandstorms in North China makes the air pollution situation more complex than ever before. The concentration of PM$_{10}$ (particulate matter with an aerodynamic diameter of less than 10 µm) in Beijing reached up to 3600 µg m$^{-3}$, largely beyond the standards of the World Health Organization and the Chinese government (G. Li et al., 2021). As the most abundant primary aerosol in the troposphere (Textor et al., 2006; Tang et al., 2016), dust particles could transport more than one full circuit around the globe within ~2 weeks (Uno et al., 2009) and concurrently participate in an array of atmospheric reactions. Heterogeneous reactions over the dust surface consume and produce various trace gases, thereby affecting the dust property and tropospheric oxidation capacity (Tang et al., 2016). One of the most extensive concerns is that the numerous surface sites of wind-blown dust provide opportunities for a variety of atmospheric reactions to occur, e.g. oxidation of SO$_2$ and formation of sulfate (Usher et al., 2003). In the past decades, plenty of laboratory works have been performed to explore the heterogeneous behaviours of SO$_2$ on dust surfaces.

When discussing the heterogeneous oxidation on a dust proxy, environmental factors like humidity, temperature, and irradiance were frequently concerned. Adsorbed water not only helps to accumulate surface S(V) species, but also competes with SO$_2$ for surface sites (Rubasinghe and Grassian, 2013). The reversible adsorption of SO$_2$ is believed to be exothermic (Clegg and Abbatt, 2001a). In contrast, there were positive temperature dependences observed for CaCO$_3$ below 250 K over the entire reaction (Wu et al., 2011) and for Fe$_2$O$_3$ within 284–318 K during the initial reaction stage (T. Wang et al., 2018a). Light irradiation normally accelerates the transformation of (bi)sulfites to (bi)sulfates (J. Li et al., 2010; Nanayakkara et al., 2012; Han et al., 2021), while iron oxides may undergo photo-reactive dissolution and thus exhibit weaker reactivity under sunlight (Fu et al., 2009). The reactivity can be enhanced by H$_2$O$_2$ (Huang et al., 2016), O$_3$ (L. Li et al., 2006, 2007; Y. Zhang et al., 2018), NO$_3$ (Ma et al., 2008; C. Liu et al., 2012; He et al., 2014; Yu et al., 2018; Zhao et al., 2018; W. Yang et al., 2018a; H. Wang et al., 2020), NH$_3$ (W. Yang et al., 2016, 2018b, 2019), or Cl$_2$ (Huang et al., 2017). By contrast, organic compounds, like CH$_2$CHO (Zhao et al., 2015), HCOOH (Wu et al., 2013), CH$_2$COOH (N. Yang et al., 2020; R. Wang et al., 2022), and C$_3$H$_6$ (Chu et al., 2019), were found to suppress the interactions due to the accumulation or production of particle-phase organic acids. Moreover, CO$_2$ hinders the heterogeneous oxidation under dark conditions (Y. Liu et al., 2020) but presents positive impacts on sulfate formation under solar irradiation (Y. Liu et al., 2022). The particle physical properties, including size (Baltrusaitis et al., 2010; Y. Zhang et al., 2016), morphology (K. Li et al., 2019), and crystal structure (W. Yang et al., 2017), display varied impacts on heterogeneous reaction. When considering the particle chemical properties, Fe$_2$O$_3$ is more active than Al$_2$O$_3$ and SiO$_2$ in the heterogeneous uptake of SO$_2$ and formation of sulfate (Chughtai et al., 1993; X. Zhang et al., 2006; He et al., 2014). Furthermore, the presence of moderate nitrate (Kong et al., 2014; Du et al., 2019), Al$_2$O$_3$ (T. Wang et al., 2018b), surfactant (Zhankazkova et al., 2019), oxalate (K. Li et al., 2021), or (bi)carbonate (Y. Liu et al., 2022) on dust surfaces could favour the heterogeneous kinetics under specific conditions. Relative to the environmental factors, the heterogeneous effects relevant to particle properties are still under debate.

The usage of a simple mineral oxide as a substitute for natural dust may be problematic, as such an approach could undermine the atmospheric importance of more complex mineralogy. Authentic dusts were utilized in laboratory works (Table S2). Some studies concerned single samples like Saharan dust (Ullerstam et al., 2002, 2003; Adams et al., 2005; Harris et al., 2012), Arizona test dust (Park and Jang, 2016; S. Zhang et al., 2019; Y. Zhang et al., 2019), Chinese loess (Usher et al., 2002), and Asian dust (Ma et al., 2012). The comparison of diverse samples has been attracting increasing...
attention. Zhou et al. (2014) observed the positive temperature effect on Xinjiang sierozem, in contrast to the negative temperature dependence for Inner Mongolian desert dust. Huang et al. (2015) discovered the accelerated oxidation of SO$_2$ by H$_2$O$_2$ and attributed the different moisture effects to the dusts’ varying components. Mørner et al. (2017) explored the heterogeneous uptake on volcanic ash and glass samples and related the reactivity differences to the varying abundances of surface basic and reducing sites. Park et al. (2017) compared the heterogeneous reactions on Gobi desert dust and Arizona test dust and linked the sulfate formation to the quantity of semi-conductive metals. Z. Wang et al. (2019) discovered the enhanced uptake of SO$_2$ on clay minerals after the simulated cloud processing and explained this evolution by the modification of iron speciation. Uropolis et al. (2019) discussed the kinetics of diverse volcanic dusts and further experimentally proved that neither one selected pristine oxide nor a mixture of them can adequately typify the behaviour of natural dust (Uropolis et al., 2021). Recently, based on the measurement method for sulfite and sulfate (Uropolis et al., 2020), they determined the associations between sulfate production and dust chemical properties like Fe + Al / Si and Na abundance (Uropolis et al., 2022). The aforementioned works broaden the horizons of the heterogeneous drivers on the dust surface. Up to now, the dominant dust surface drivers have remained controversial due to the limited statistical linkages between the chemical composition of dust and the production rate of sulfate.

While the atmospheric relevance of the oxidation of SO$_2$ on dust has been widely recognized (Wu et al., 2020; Xu et al., 2020), the contribution of dust heterogeneous oxidation to secondary sulfate aerosols has not been quantitatively determined. By means of the improved WRF-CMAQ model, K. Wang et al. (2012) attributed a 27% decrement of SO$_2$ concentration and a 12% increment of sulfate concentration to the heterogeneous processes during an Asian dust storm. Moreover, Tian et al. (2021) recently simulated the heterogeneous formation of dust sulfate with the revised GEOS-Chem model and found that, during the dust episodes in North China, up to 30% of the secondary sulfate resulted from the heterogeneous processes on the dust surface. However, the photocatalytic reactivity of dust was not considered by the advanced models. Yu et al. (2017) developed the atmospheric mineral aerosol reaction (AMAR) model based on the laboratory works and suggested that the heterogeneous photocatalysis of the mineral dust surface contributed more than half of the secondary sulfate. However, the heterogeneous reactivity was measured by quantifying all the adsorbed SO$_2$ rather than calculating the yield of particle-phase sulfate. Moreover, limited gas- and aqueous-phase pathways were included in these models, and the heterogeneous reactivities of surface-adsorbed oxidants and dust surface drivers have not been distinguished yet. Therefore, it is highly desirable to comprehensively compare the dust heterogeneous pathways with other documented sulfate formation pathways.

Hereby, upon understanding the driving factors and driving force of the airborne dust surface, this work compared dust heterogeneous pathways with the gas- and aqueous-phase ones with respect to the formation rate of sulfate and the atmospheric lifetime of SO$_2$. In order to characterize the sensitivity of heterogeneous reaction to dust loading, the scenarios with different dust concentrations were also considered. The joint influences of ionic strength and aerosol liquid water content on the aqueous-phase oxidation of SO$_2$ were further discussed to prove the significance of heterogeneous kinetics under diverse atmospheric conditions. The recently reported microdroplet interfacial oxidations were additionally compared with the dust-related heterogeneous pathways to emphasize the atmospheric relevance of dust surface drivers. This study attempted to verify the significant formation of sulfate aerosols contributed by the heterogeneous drivers of the dust surface.

## 2 Methods

### 2.1 Methodology overview

This study attempted to investigate whether the heterogeneous oxidation of SO$_2$ on the dust surface, particularly that induced by the dust surface drivers, makes great impacts on the loss of gaseous SO$_2$ and formation of sulfate aerosols (Text S1, Fig. S1). The gas- and aqueous-phase pathways were assessed by the documented methodologies and parameterizations (Cheng et al., 2016; Seinfeld and Pandis, 2016; Shao et al., 2019; Song et al., 2021), as briefly introduced by Sect. 2.2 (more details in the Supplement). The heterogeneous conversion of SO$_2$ comprises dust-mediated and dust-driven modes, emphasizing the oxidants co-adsorbed with SO$_2$ and the drivers of the dust surface, respectively. In the dust-mediated mode, the dust surface functions as a reaction medium that supports the interaction between adsorbed oxidants and SO$_2$. In the dust-driven mode, the oxidation of adsorbed SO$_2$ is initiated by the active components of the dust surface. The former mode was assessed by the particle properties based on the reported methodology, while the latter mode was quantitatively characterized by the laboratory works.

### 2.2 Gas- and aqueous-phase oxidation pathways

The gas-phase oxidation of SO$_2$ is initiated by a hydroxyl radical (OH), stabilized Criegee intermediates (CIs), and a nitrate radical (NO$_3$). The former two oxidants promote the sulfate formation during daytime, while the latter one works mainly during nighttime. More details can be found in Text S2.

The aqueous-phase sulfate formation is pH-dependent and can be quantified based on the published documents and the references cited therein (Cheng et al., 2016; Su et al., 2020; T. Liu et al., 2021a). Herein, an aqueous-phase pathway
refers to the liquid SO$_2$ conversion by the transition-metal ion-catalysed oxygen (TMI-O$_2$), ozone (O$_3$), hydrogen peroxide (H$_2$O$_2$), nitrogen oxide (NO$_2$), methyl hydroperoxide (CH$_3$OOH), peroxyacetic acid (CH$_3$COOOH), hypohalous acid (HOCI), hypobromous acid (HOBr), dissolved nitrous acid (HONO), photosensitization ($T^*$), and nitrate photolysis ($P_{NO_3}^-$). The processes relevant to $T^*$ and $P_{NO_3}^-$ are only considered for the daytime scenario. More details can be found in Text S3 and Tables S3–S4.

The oxidant concentration data were derived from the atmospheric observation campaigns performed in Beijing, North China. The measurements conducted in warm seasons were considered in priority to correspond to the experimental temperature. The relevant temperature range is 293–303 K and is representative of the warm season. Considering the relatively high irradiance used in the laboratory experiments, the oxidant parameters for daytime discussion were selected from the observations performed at noontime. The influences of dust loading on the oxidant concentrations, nitrate photolysis kinetics, and TMI abundances were considered to reflect the linkages between different reactions. More details can be found in Text S4 and Table S5.

### 2.3 Heterogeneous oxidation–particle characterizations

Five airborne clay minerals, including nontronite (NAu), chlorite (CCa), montmorillonite (SWy), kaolinite (KGa), and illite (IMt), were obtained from the Source Clay Minerals Repository (Purdue University, West Lafayette, Indiana, USA). The purchased clays were sent for the following measurements.

The clay minerals were analysed with an X-ray fluorescence spectrometer (Axios Advanced, PANalytical, Netherlands) for element distributions (Table S6). The Brunauer–Emmett–Teller (BET)-specific surface areas ($S_{BET}$) of NAu, CCa, SWy, KGa, and IMt were measured with a Quantachrome Nova 1200 BET apparatus to be 19.76, 5.67, 22.64, 18.77, and 20.05 m$^2$ g$^{-1}$, respectively.

Particles were suspended in Milli-Q water (18.2 MΩ cm at 25°C) before the size measurement (ViewSizer$^\text{TM}$ 3000, MANTA Instruments, USA). The particle diameters range mostly from 50 to 1000 nm and are averaged to 399 nm for NAu, 272 nm for CCa, 438 nm for SWy, 396 nm for KGa, and 366 nm for IMt (Fig. S2).

Prepared particles were ultrasonically extracted in Milli-Q water, followed by the filtration through a polytetrafluoroethylene membrane filter. The obtained solution was analysed by ion chromatography (883 Basic, Metrohm, Switzerland) for the concentrations of anions (HCOO$^-$, Cl$^-$, NO$_3^-$, SO$_4^{2-}$) and cations (Na$^+$, NH$_4^+$, K$^+$, Mg$^{2+}$, Ca$^{2+}$) (Fig. S3) by the reported methods (Wang et al., 2020b). The watersoluble ions account for 4.6%, 0.1%, 8.5%, 0.3%, and 0.8% of the mass contents of NAu, CCa, SWy, KGa, and IMt, respectively.

### 2.4 Heterogeneous oxidation–DRIFTS measurements

The in situ DRIFTS (diffuse reflectance infrared Fourier transform spectroscopy) spectra were collected using a FTIR spectrometer (Tracer-100, Shimadzu, Japan) equipped with a mercury–cadmium–telluride detector cooled by liquid nitrogen. A gas supply system was constructed by linking the experimental units through Teflon tubes. Mass flow controllers (D07-19, Severstar, China) adjusted the reactant gases to the expected reactant concentration and relative humidity (RH). Gas cylinders had high-purity air (79% N$_2$ and 21% O$_2$) and SO$_2$ (2.46 x 10$^{15}$ molecules cm$^{-2}$ diluted by N$_2$).

Before each experiment, a ceramic cup holding particles was placed in the reaction chamber. The particles were treated in a stream of dry air (300 mL min$^{-1}$) for 30 min to minimize the surface water and impurities (T. Wang et al., 2018a, c). After the pretreatment, the sample was exposed to humidified air (RH = 50%) to reach moisture saturation, followed by the collection of the background spectrum and then the introduction of a reactant gas for 240 min. The SO$_2$ concentration was 3.69 x 10$^{13}$ molecules cm$^{-2}$ in a total flow rate of 100 mL min$^{-1}$.

The simulated solar irradiation with an actinic flux of 6.51 x 10$^{15}$ photons cm$^{-2}$ s$^{-1}$ was provided by a xenon lamp (TCX-250, Ceaulight, China). The reaction temperature (296.8 K) was controlled by a heater attached to a recirculating cooling water system and determined by a calibration curve introduced previously (T. Wang et al., 2018a). The kinetic parameters involved in the gas- and aqueous-phase oxidation pathways (listed by Sect. 2.2) were corrected by the experimental temperature as much as possible. All exposures were performed in triplicate. The experimental set-ups are displayed by Fig. S4, and the recorded spectra are shown in Figs. S5 and S6.

### 2.5 Heterogeneous oxidation – acidity and kinetics

The recorded spectra were analysed by referring to the previous literature and the references cited therein (Persson and Vgren, 1996; Peak et al., 1999; Goodman et al., 2001; X. Zhang et al., 2006; Wu et al., 2011; C. Liu et al., 2012; Nanayakkara et al., 2012, 2014; Huang et al., 2016; Ma et al., 2017; W. Yang et al., 2017). In total, six sulfur-containing species can be identified: hydrated SO$_2$ (SO$_2$·H$_2$O), bisulfite, sulfite, solvated sulfate, coordinated

A mixture, denoted as natural dust hereafter, was prepared by mechanically mixing the studied clay minerals with their atmospheric abundances. Because the clays in the kaolinite group (NAu, KGa), montmorillonite group (SWy), illite group (IMt), and chlorite group (CCa) occupy, respectively, 6.6%, 4.0%, 53.8%, and 4.3% mass fractions of the airborne dust (Usher et al., 2003), the prepared natural dust sample comprises 8.8 wt% of NAu, 8.8 wt% of KGa, 5.3 wt% of CCa, 71.4 wt% of IMt, and 5.7 wt% of CCa.
sulfate, and bisulfate, and the assignments are summarized by Table S7. The overlapping bands were further analysed by Gaussian/Lorentzian deconvolution to obtain the product distributions (Fig. S7). The consistent deconvolution procedure was performed for the infrared spectra derived from the repeated experiments.

Because some of the measured ions exist in the surface-coordinated forms or crystalline states, the particle acidity (pH) may not be accurately characterized by the typical proxy methods (Hennigan et al., 2015). Herein, the ionization equilibrium of dissolved SO$_2$ in the water layers of the particle surface is considered to be associated with particle acidity. As reported, the dissolved SO$_2$ would transform from SO$_2$·H$_2$O to HSO$_3^-$ and then to SO$_3^{2-}$ as the medium evolves from the extremely acidic to the nearly alkaline (Haynes, 2014; R. Zhang et al., 2015). The relative abundance of SO$_2$·H$_2$O and SO$_3^{2-}$ as assumed to be equivalent to the relative integral area of their characteristic peaks, is utilized to calculate the particle acidity (more details in Text S3–S2 of the Supplement). The relative abundance of HSO$_4^-$ and SO$_2^-$, which can be used to calculate the acidity (Rindelaub et al., 2016; Ault, 2020), was not considered by the present study because the characteristic signals of bisulfate cannot be observed in some of the recorded spectra. As noted by Keene et al. (2004) and Hennigan et al. (2015), estimating the pH of airborne aerosol by ion balance may be difficult because it is unable to distinguish between free and undissociated H$^+$ (e.g. protons associated with HSO$_4^-$ and HSO$_3^-$). Herein, the relative abundance of $S_{(IV)}$ species is derived from the different infrared absorption signals relevant to atomic sulfur. Thus, the current method is recommended for the pH determination of the humidified gas–solid interface.

Heterogeneous kinetics can be assessed by the reactive uptake coefficient ($\gamma$) by assuming a first-order loss of SO$_2$. The $\gamma$ for dust-driven heterogeneous reactions can be calculated by

$$ \gamma = \frac{d[\text{SO}_4^{2-}]/dt}{Z}, $$

where $d[\text{SO}_4^{2-}]/dt$ is the rate of sulfate production on the dust surface (ion s$^{-1}$), $A_s$ is the reactive surface area (m$^2$), $[\text{SO}_2]$ is the experimental concentration of SO$_2$ (molecules m$^{-3}$), $v_{\text{SO}_2}$ is the molecular velocity of SO$_2$ (m s$^{-1}$), $R$ is the gas constant (J mol$^{-1}$ K$^{-1}$), $T$ is the experimental temperature (K), and $M_{\text{SO}_2}$ is the molecular weight of SO$_2$ (kg mol$^{-1}$). Because the infrared intensity is proportional to the amount of surface product, $[\text{SO}_4^{2-}]$ can be translated by the integral area of the sulfate characteristic peaks:

$$ [\text{SO}_4^{2-}] = f \times (\text{integral area}), $$

where $f$ is the conversion factor and represents the amount of SO$_4^{2-}$ corresponding to the per-unit integral area. The sulfate production rate of the dust surface can be translated by the calibration curves by mixing weighed Na$_2$SO$_4$ with the target particle sample to a set of concentrations (Martin et al., 1987; L. Li et al., 2006; Wù et al., 2011; T. Wang et al., 2020a). The conversion factors are calculated as $1.32 \times 10^{18}$, $4.62 \times 10^{17}$, $6.97 \times 10^{17}$, $8.20 \times 10^{17}$, $9.44 \times 10^{17}$, and $9.25 \times 10^{17}$ for NAu, CCa, SWy, KGa, IMt, and natural dust, respectively. Because Na$_2$SO$_4$ was thoroughly mixed with the particles, $\delta_{\text{BET}}$ was used to calculate $\gamma$.

All samples except SWy presented steady sulfate production potentials over the entire experiment, while the products in SWy increased by reaction time in the beginning and then gradually remained unchanged (Fig. S8). Accordingly, the process in SWy was assessed by the experimental data of the first 30 min of the reaction, and the kinetics of other samples were calculated by the spectra recorded throughout the experiment.

Besides the dust surface drivers, the heterogeneous conversion of SO$_2$ can be additionally initiated by the gaseous oxidants (O$_3$, H$_2$O$_2$, NO$_2$, HOCl, HOB or OH, CH$_3$OOH, CH$_3$COOH, HONO) co-adsorbed with SO$_2$. Considering that the adsorption of the oxidant onto dust surfaces would produce gas- and particle-phase species (Usher et al., 2003; Tang et al., 2017), the dust-mediated heterogeneous oxidation can be assumed to primarily proceed in the surface water layers. Such an assumption can be experimentally proven as the acceleration of SO$_2$ oxidation induced by the co-oxidant (e.g. H$_2$O$_2$, NO$_2$, O$_3$) becomes more significant under higher humidity (Huang et al., 2015; Park et al., 2017; Y. Zhang et al., 2018). Based on the literature (Hanson et al., 1994; Jacob, 2000; Seinfeld and Pandis, 2016; Shao et al., 2019), the $\gamma$ for dust-mediated heterogeneous reaction can be calculated by

$$ \gamma = 1 - \frac{1}{\alpha} - \frac{v}{4H^*RT\sqrt{D_\text{chem}}}, $$

$$ k_{\text{chem}} = \frac{R_a}{\sum S_{(IV)}}, $$

$$ f_t = \frac{\coth r_p}{l} - \frac{l}{r_p}, $$

$$ l = \sqrt{\frac{D_a}{k_{\text{chem}}}}, $$

where $k_{\text{chem}}$ is the pseudo first-order reaction rate constant of the studied $S_{(IV)}$ species (s$^{-1}$) that is produced at a rate of $R_a$ (M s$^{-1}$), $f_t$ is the diffusive correction term comparing the radius of natural dust $r_p$ (m) with the diffuse-reactive length $l$ (m), $\alpha$ is the mass accommodation coefficient of SO$_2$ (dimensionless), $D_a$ is the aqueous-phase diffusion coefficient.
of SO$_2$ (1.78 $\times$ 10$^{-5}$ m$^2$ s$^{-1}$ at 296.8 K) (Himmelblau, 1964; Haynes, 2014), $H^*$ is the effective Henry’s law constant for the studied $S_{(IV)}$ species (M atm$^{-1}$) and is jointly determined by the gas–liquid equilibrium of SO$_2$ and the ionization equilibria of dissolved sulfur species, and $R$ is the gas constant (L atm mol$^{-1}$ K$^{-1}$). The terms on the right-hand side (RHS) of Eq. (5) show the two contributions to the overall resistance to dust-mediated heterogeneous uptake: mass accommodation at the surface, diffusion, and reaction in surface liquid water layers. The $S_{(IV)}$ species include SO$_2$·H$_2$O, HSO$_3^-$, and SO$_4^{2-}$.

The sulfate formation rate in the atmosphere can be calculated by the following equations (Jacob, 2000; J. Li et al., 2020).

$$\frac{d [SO_4^{2-}]}{dt} = \left( \frac{r_p}{D_g} + \frac{4}{v}\gamma \right)^{-1} S_p [SO_2], \quad (9)$$

$$S_p = C \times F \times S_{\text{BET}}, \quad (10)$$

where $d [SO_4^{2-}] / dt$ is the atmospheric sulfate formation rate (µg m$^{-2}$ h$^{-1}$), $r_p$ is the particle radius (m), $D_g$ is the gas-phase diffusion coefficient of SO$_2$ (m$^2$ h$^{-1}$), $v$ is the molecular velocity of SO$_2$ (m s$^{-1}$), $\gamma$ is the reactive uptake coefficient (dimensionless), $S_p$ is the particle surface area density (m$^2$ m$^{-3}$), $[SO_2]$ is the atmospheric SO$_2$ concentration (µg m$^{-3}$), $C$ is the dust concentration of 55 µg m$^{-3}$ representative of the common atmospheric condition of China (X. Y. Zhang et al., 2012), $F$ is the mass fraction of clay minerals in the natural dust community (%), and $S_{\text{BET}}$ is the BET-specific surface area (m$^2$ g$^{-1}$).

2.6 Atmospheric lifetime of SO$_2$

For the gas- and aqueous-phase pathways, the lifetime of SO$_2$ ($\tau$) can be calculated by Eq. (11) (Jacob, 2000).

$$\tau = \frac{1}{k_{\text{chem}}}, \quad (11)$$

where $k_{\text{chem}}$ is the assumed first-order reaction rate constant of the studied $S_{(IV)}$ species (s$^{-1}$).

For the heterogeneous pathways, the $\tau$ can be calculated by Eq. (12) (Clegg and Abbatt, 2001b).

$$\tau = \frac{4}{\gamma v S_p}, \quad (12)$$

where $\gamma$ is the reactive uptake coefficient (dimensionless), $v$ is the molecular velocity of SO$_2$ (m s$^{-1}$), and $S_p$ is the particle surface area density (m$^2$ m$^{-3}$), as described by Eq. (10).

The atmospheric lifetime caused by the multiple pathways ($\tau_{\text{total}}$) can be estimated by Eq. (13) (Seinfeld and Pandis, 2016).

$$\tau_{\text{total}} = \left( \sum \frac{1}{\tau} \right)^{-1} \quad (13)$$

3 Results and discussion

3.1 Driving factors of the dust surface

Correlation analysis is performed to identify the dust surface drivers (Fig. 1), based on which the reaction mechanism of a dust heterogeneous pathway can be better understood (Fig. 2).

The mineral drivers are investigated first (Fig. 1a). Under dark conditions, the sulfate production rate correlates positively with Fe but presents negative dependence against Al. As documented, airborne sulfate was highly associated with Fe-rich dusts, and the heterogeneous reaction of SO$_2$ can be regarded as a possible explanation (Sullivan et al., 2007). Typical Fe-bearing minerals accelerate the oxidation of $S_{(IV)}$ species by either the surface-active oxygen (O$^-$) derived from the adsorbed O$_2$ in an oxygen vacancy or the iron redox cycling initiated by Fe$^{3+}$ from surface acidic media or both (Baltrusaitis et al., 2007; Fu et al., 2007; W. Yang et al., 2016). Conversely, Al$_2$O$_3$ presents weaker heterogeneous reactivity than Fe$_2$O$_3$ (X. Zhang et al., 2006; W. Yang et al., 2018b; Xu et al., 2021) and may hinder heterogeneous reaction by blocking the active sites of other mineral constituents (T. Wang et al., 2018b). Generally, an Fe-bearing component plays a crucial role in the heterogeneous reaction of SO$_2$, while the presence of an Al-bearing component directly weakens the dust’s reactivity or indirectly decreases the proportion of other active mineral constituents.

The sulfate yield enhanced by solar irradiation associates positively with the abundance of Ti or Al. Transition metal oxides in dust act as a photocatalyst that yields electron–hole pairs, followed by the formation of reactive oxygen species (ROS) such as hydroxyl radicals (·OH), superoxides (O$_2^-$), hydroperoxyl radicals (HO$_2^-$), and dissociated active oxygen species (O$^+$) (Chen et al., 2012; Abou-Ghanem et al., 2020; T. Wang et al., 2020b; Sakata et al., 2021). Apart from the site-blocking effect under dark conditions, irradiated Al-bearing constituents additionally facilitate the sulfate formation. Physically, Al-bearing components disperse other efficient mineralogical constituents in the case of agglomerated Al-bearing constituents additionally facilitate the sulfate formation process (Guan et al., 2014). Generally, the kinetic discrepancy between dark reactions and photo-reactions relates primarily to the existence of Ti- and Al-bearing minerals.

No correlation can be observed between the abundance of elements and the sulfate production rate of photo-reaction. The dust with a higher proportion of elemental Ti was reported to exhibit greater reactivity toward SO$_2$ (Park et al., 2017), NO$_2$ (Ndour et al., 2009), or O$_3$ (Abou-Ghanem et al., 2020) in the photochemical processes. However, these results were derived from qualitative comparisons rather than the quantitative analysis performed herein. Analogous to the Al$_2$O$_3$ discussed above, semi-conductive metal exhibits dual
behaviours in dark and light reactions. For instance, the iron oxide in dark reactions has great reactivity toward SO$_2$, whereas that under solar irradiation may present impaired heterogeneous kinetics due to dissolution (Fu et al., 2009). In addition, the heterogeneous reaction on the surface of TiO$_2$- or Ti-bearing mineral is thermodynamically favoured under dark conditions but depends largely on photocatalysis under solar irradiation (Chen et al., 2012). Due to the complexity of photo-reaction that involves both dark and light reaction mechanisms, the sulfate formation rate of photo-reaction may not be directly linked to particle chemical compositions.

The driving effects of water-soluble ions are further studied (Fig. 1b). Herein, Na$^+$ and Cl$^-$ are observed to present positive impacts on the sulfate formation under dark conditions. The presence of halite (NaCl) has positive implications for the dust’s hydroscopic growth (Y. Wang et al., 2014; Tang et al., 2019), and the moderate surface water provides an efficient medium for the adsorption of gas-phase SO$_2$ and the ionization of dissolved sulfur species (Rubasinghe and Grassian, 2013). Moreover, there are negative associations between the photo-induced sulfate enhancement and the abundances of Na$^+$, K$^+$, and Ca$^{2+}$. These cations can be hydrolysed by adsorbed water to produce H$^+$, thereby elevating the particle acidity. In the surface aqueous medium, increased acidity retards the hydrolysis and dissociation of SO$_2$ (Park et al., 2008; Huang et al., 2015) and inhibits the production of OA by the irradiated dust surface (Zheng et al., 1997; L. Yang et al., 2008; X. Liu et al., 2017). Urupina et al. (2022) reported the positive correlations between secondary dust sulfate and the amount of elemental Na. The results here provide additional insights into the heterogeneous effects of the water-soluble constituents on the dust surface.

Apart from the dust surface drivers and inhibitors, the sulfate radical (SO$_4^{2-}$) may also participate in the heterogeneous event. In the absence of solar irradiation, laboratory studies discovered the non-catalysed SO$_2$ autoxidation by SO$_4^{2-}$ at the acidic droplet interface (Hung and Hoffmann, 2015; Hung et al., 2018; Chen et al., 2022). With an appropriate humidity and sufficient saline component, there could be an air–liquid interface occurring over the solid dust surface via water uptake (Gaston et al., 2017; Tang et al., 2019; Wu et al., 2020), producing SO$_4^{2-}$. Moreover, on the surface of iron oxide, the $^\cdot$OH derived from a Fenton reaction reacts with SO$_4^{2-}$ to form SO$_3^{2-}$ (Kim et al., 2019; T. Li et al., 2020). The presence of solar irradiation would accelerate the formation of SO$_3^{2-}$ due to the abundant $^\cdot$OH produced via photocatalysis (Antoniou et al., 2018). Generally, the dust surface with a higher proportion of saline components and transition metals may cause the more significant oxidation by SO$_4^{2-}$, especially under solar irradiation with higher ambient humidity and particle acidity.

Based on the correlation analysis, regression analysis can be further performed to predict the surface kinetics by the chemical compositions of dust. Referring to the publication of F. Zhang et al. (2020), we derived an exponential parameterization for the production rate of particle-phase sulfate:


where $S_{dark}$ is the yield of SO$_4^{2-}$ on the particle sample of per-unit mass within per-unit time (ions g$^{-1}$ s$^{-1}$). $[A], [B], [C], [D], [E]$, and $[F]$ are the mass fractions of the element and ions (%). $M$ and $N$ are the constant parameters (dimensionless).

After the statistical procedures by SPSS (version 22.0), we obtained the regression equation for the dark reaction:

$$S_{dark} \times 10^{15} = 22.858 [AI]^{-0.001} [Fe]^{0.111} [Cl^-]^{-0.001} [Na^+]^{0.001} - 13.404.$$  

(15)

Analogously, the discrepancy between the sulfate production rates under dark and illuminated conditions, denoted as $R_{light/dark}$, can be described by

$$R_{light} = S_{light}/S_{dark} = 14.539 [AI]^{0.073} [Ti]^{0.167} [Cl^-]^{-0.001} [Na^+]^{-0.001} [K^+]^{-0.001} [Mg^{2+}]^{-0.001} - 3.186.$$  

(16)

Then,

$$S_{light} = S_{dark} \times R_{light/dark}.$$  

(17)

Because the abundance variation of a water-soluble ion presents significantly less influence on the prediction result relative to the elemental transition metal, the complete regression models can be simplified by merely considering the mineral element abundances:

$$S_{dark} \times 10^{15} = 30.880 [Fe]^{0.077} - 21.679,$$

$$R_{light/dark} = 15.581 [AI]^{0.062} [Ti]^{0.141} - 4.277,$$

$$S_{light} = S_{dark} \times R_{light/dark}.$$  

(19)

The complete and simplified regression models both accurately simulate the experimental data, with all linear slopes approaching 1.0 and all $R^2$ values larger than 0.995 (Fig. 3). The simplified regression model can be considered the preferred recommendation due to the fewer parameters needed and its greater performance in the photo-reaction prediction. Shang et al. (2010) found that the heterogeneous sulfate production on pristine TiO$_2$ (Degussa P25) can be accelerated by 8.4 times by the presence of ultraviolet light (365 nm, 350 µW cm$^{-2}$). Such photo-induced enhancement is comparable with the prediction result (10.2 times) derived from the simplified regression model developed by this study. Additionally, because the chemical composition of a bulk sample may not fully explain the sulfate formation over the dust surface, further study to discuss the model uncertainty is warranted.
3.2 Driving force of the dust surface

The driving force of the dust surface can be characterized by a reactive uptake coefficient and particle acidity (Fig. 4, Table S8). Under dark conditions, the $\gamma$ values are highest for SWy and NAu, followed by IMt and CCa, with that of KGa being lowest. The presence of simulated solar irradiation causes a different rank: IMt $>$ SWy $>$ CCa $>$ NAu $>$ KGa. The $\gamma$ values of CCa, SWy, KGa, and IMt under dark conditions are, respectively, 1.52, 1.01, 2.94, and 2.30 times greater than those for photo-reactions, reflecting the distinct photocatalytic performances of the clay minerals. Conversely, NAu presents the decreased heterogeneous uptake capacity when exposed to solar irradiation due to its rich abundance of Fe, whose oxides may occur in photo-reductive dissolution in acidic media (Fu et al., 2010; Shi et al., 2012).

Comparing the studies performed by the same experimental approach and assessment procedures, these clay minerals are more efficient than the previously concerned mineral dust proxies, including CaCO$_3$ (L. Li et al., 2006; Wu et al., 2011; Y. Zhang et al., 2018), Al$_2$O$_3$ (W. Liu et al., 2017), TiO$_2$ (He and Zhang, 2019), and manganese oxides (T. Wang et al., 2020a) in the heterogeneous production of sulfate under the parallel conditions.

The $\gamma$ values for the natural dust-driven heterogeneous reactions are calculated as $6.08 \times 10^{-6}$ under dark conditions and $1.14 \times 10^{-5}$ under illuminated conditions (Fig. 4a). Within the dataset of authentic dust, the $\gamma$ for dark reactions is lower than the reported uptake coefficients of Chinese loess ($3.0 \times 10^{-5}$), Inner Mongolian desert dust ($2.41 \times 10^{-5}$), Xinjiang sierozem ($8.34 \times 10^{-5}$), Saharan dust ($6.6 \times 10^{-5}$), Asian mineral dust ($2.54 \times 10^{-5}$), Tengger desert dust ($4.48 \times 10^{-5}$), and Arizona test dust (ATD) ($1.92 \times 10^{-5}$) under the similar experimental conditions (Usher et al., 2002; Adams et al., 2005; Zhou et al., 2014; Huang et al., 2015). It should be noted that the previous studies measured the net uptake coefficient that quantifies all the heterogeneously adsorbed SO$_2$, and some of them assessed the oxidations accelerated by surface-adsorbed oxidants (e.g. NO$_2$, O$_3$, or H$_2$O$_2$). In order to quantify the driving force of the dust surface in the contribution of particle-phase sulfate, $S_{(IV)}$ species were not involved in the current kinetics calculation. Quantitatively, the dust surface drivers are responsible for the atmospheric sulfate formation rates of 0.195 and 0.365 µg m$^{-3}$ h$^{-1}$ during nighttime and daytime, respectively.

Particle acidity is further calculated to discuss the driving force of the dust surface for sulfate formation. After the SO$_2$ exposure, CCa is the most acidic one, followed by IMt, KGa, and NAu, leaving SWy more neutral. Because no significant correlation can be found between the $\gamma$ for dust-driven sulfate formation and the pH of reacted clay mineral, the absolute acidity level depends largely on the basic nature of dust. The lowest pH assigned to CCa can be explained by its highest content of elemental Mg relative to the other clays (see Table S6). The Mg-bearing constituents dissolve to Mg$^{2+}$, which can be further hydrolysed by water to produce H$^+$, thus accelerating the acidification of particles (Park et al.,

**Figure 1.** Correlation analysis of the sulfate production rate of the dust surface. Spearman correlation coefficient for the relationship between the rate of sulfate production and the abundance of (a) mineral elements or (b) water-soluble ions. The significant correlations with $p$ values lower than 0.05 or 0.01 are highlighted by the labels in white.
Figure 2. Mechanism illustration of the dust-driven heterogeneous conversion of SO₂. The heterogeneous reaction of airborne SO₂ on a natural dust surface includes the initial adsorption of gas-phase SO₂ and the later conversion of S(IV) species to S(VI) products. The formed S(IV) species include both adsorbed and dissolved SO₂, bisulfite, and sulfite (dashed box in blue). The generation of S(VI) products can be attributed to the surface active oxygen (O²⁻), Fe(III)–Fe(II) redox cycling, reactive oxygen species (ROS), and sulfate radical (SO₄²⁻) (dotted box in deep yellow). Heterogeneous oxidation is largely influenced by the dust surface drivers and inhibitors, which promote and hinder the heterogeneous procedures, respectively (solid box in black). Dust surface drivers are determined as (a) the Fe-bearing constituents and saline ions (Na⁺, Cl⁻) under dark conditions and (b) the Ti- and Al-bearing constituents under solar irradiation. Dust surface inhibitors are identified as (c) the Al-bearing constituents under dark conditions and (d) the metal cations (Na⁺, K⁺, Ca²⁺) under solar irradiation. Figure 5 presents the pH-dependent γ for the natural dust-mediated heterogeneous pathway along with the experimental data of the natural dust-driven heterogeneous pathway. The dust-mediated γ is orders of magnitude lower than the α of SO₂ (∼0.14 under the experimental temperature of 296.8 K). By definition, α is the probability that a SO₂ molecule striking a liquid surface will finally enter the liquid phase, whereas γ is the sulfate formation rate normalized by the total surface collision rate of SO₂ [Jacob, 2000; Davidovits et al., 2006]. That is, γ involves all the uptake processes, including the mass accommodation at the surface (Seinfeld and Pandis, 2016). Moreover, unlike the pH-independent α, γ varies with pH, somewhat coinciding with the evolution of dissolved sulfur species that can be quantitatively described by the H⁺ of SO₂. The γ values of H₂O₂, CH₃OOH, and CH₃COOOH are pH-independent under spe-
Specific conditions in relation to the acid-catalysed rate-limiting steps relevant to these peroxides (Lind et al., 1987; T. Liu et al., 2021). Generally, the dust-mediated heterogeneous γ is largely determined by the diffusion and reaction processes in the water layers of the dust surface, as characterized by the second item on the RHS of Eq. (5).

The dust-mediated heterogeneous sulfate formation is primarily contributed by the surface-adsorbed H₂O₂ below the nocturnal pH of 5.50 or the diurnal pH of 5.27. When the acidity exceeds the thresholds, the dust-mediated pathway would be kinetically dominated by NO₂ and O₃ during nighttime or HOBr and HOCl during daytime. Comparing the γ values with the same acidity, the dust-driven pathway appears to be more efficient than the dust-mediated one, thus having the possibility of accounting for more secondary sulfate aerosols. In the following sections, the oxidation of SO₂ mediated or driven by the natural dust would be set as the typical dust-mediated and dust-driven heterogeneous reactions to compare them with the widely documented gas- and aqueous-phase pathways.

### 3.3 Comparison of atmospheric oxidation pathways

Figure 6 compares the sulfate formation rates of diverse atmospheric oxidation pathways with a newly developed comparison model. Based on the parameterization scheme in the model and the experimental results discussed above, the total sulfate formation rates are summed to 0.795 µg m⁻³ h⁻¹ during nighttime and 5.179 µg m⁻³ h⁻¹ during daytime under the acidity of natural dust (Fig. 6a and e). The higher sulfate flux during daytime can be explained by the photo-increased oxidation channels, oxidant concentrations, heterogeneous reactivity, and the elevated particle acidity that facilitates the TMI catalysis. It is worthwhile mentioning that the nocturnal sulfate concentration was reported to approach or exceed the subsequent diurnal level, as explained by the higher nocturnal humidity facilitating the liquid oxidations or the lower boundary layer at night causing the adverse diffusion conditions (J. Liu et al., 2017; Tutsak and Koçak, 2019; S. Li et al., 2020). Meteorological factors like humidity are not considered by the current model, which emphasizes the comparison of diverse pathways through the kinetic regime. The estimated sulfate fluxes are generally lower than some published data derived from the same concentration of SO₂ (40 ppb).
Figure 4. Analysis results of the in situ infrared spectra recorded for the heterogeneous reaction of SO$_2$ on clay minerals and natural dust. (a) Reactive uptake coefficients ($\gamma$) for the heterogeneous formation of sulfate. (b) Particle acidity (pH) of the reacted particle samples. The dark (grey square) and light (yellow circle) conditions were both considered. Dots represent the results of clay minerals, and those of natural dust are showed by the lines. All error bars represent 1 SD.
Figure 5. Particle acidity-dependent reactive uptake coefficients ($\gamma$) for the dust-mediated and dust-driven heterogeneous pathways. The dust-mediated pathway can be induced by the surface oxidants co-adsorbed with SO$_2$, including hydrogen peroxide (H$_2$O$_2$), nitrogen oxide (NO$_2$), ozone (O$_3$), hypochlorous acid (HOCl), hypobromous acid (HOBr), methyl hydroperoxide (CH$_3$OOH), peroxyacetic acid (CH$_3$COOOH), and dissolved nitrous acid (HONO). The dust-driven pathway can be induced by the heterogeneous drivers (transition-metal-bearing components and water-soluble ions) on natural dust and clay minerals – NAu, CCa, SWy, KGa, and IMt. More than those of the natural dust-mediated one. Hence, the heterogeneous sulfate formation is primarily ascribed to the dust surface drivers rather than the surface-adsorbed oxidants, as experimentally proven by the laboratory studies concerning the heterogeneous SO$_2$ oxidation on authentic dusts accelerated by the presence of NO$_2$ (Park et al., 2017), O$_3$ (Park et al., 2017), or H$_2$O$_2$ (Huang et al., 2015). The kinetic discrepancy between the dust-mediated and dust-driven heterogeneous pathways is more significant during nighttime than daytime. In the estimation, H$_2$O$_2$ is the predominant dust-mediated oxidant, and its concentration becomes lower under weaker sunlight. The dust-mediated contribution during nighttime is thus relatively small by the presence of the relatively low H$_2$O$_2$ concentration. In general, when investigating the heterogeneous process on dust particles, particle variables could be more important than gas variables in elucidating the reaction characteristics and atmospheric implications.

Figure 7 summarizes the particle acidity-dependent sulfate formation rates of gas-phase, aqueous-phase, and heterogeneous pathways. Relative to the pH-independent gas-phase process, the aqueous-phase oxidation becomes relatively productive under the extremely acidic and near-neutral situations but fails to support the rapid sulfate formation within the pH range of 4–5 (minima at 4.85 for nighttime and 4.55 for daytime), which overlaps with the acidity range of the weak dust-mediated heterogeneous pathway. The significant sulfate formation in the acidic medium is primarily contributed by the TMI-O$_2$ pathway where the concentrations of TMIs and H$^+$ keep increasing as the aerosol water becomes more acidic. When pH > 5.0, the sulfate formation rate increases with elevated pH, in step with the evolution of the concentration of dissolved sulfur species. Noticeably, the pH range of 4–5 involves the acidity of the aged natural dust and overlaps with that of the common haze aerosols, implying that the dust surface drivers have profound impacts on the secondary sulfate burst in a highly polluted environment (Ding et al., 2019; Song et al., 2019). The important role of dust in sulfate formation was confirmed by atmospheric observation research. For example, secondary sulfate was observed to accumulate on the dust-dominant super-micron particles collected in the North China Plain, and the mass fraction of coarse-mode sulfate dramatically increased during the evolutionary stages of haze episodes (Xu et al., 2020). Within the entire pH range, the aqueous-phase pathway is more active than the dust-mediated heterogeneous one in contributing to secondary sulfate.

Figure 8 compares the lifetimes of SO$_2$ influenced by the diverse atmospheric oxidation pathways. Calculations of lifetimes can be useful in estimating how long the SO$_2$ is likely to remain airborne before it is removed from the atmosphere (Seinfeld and Pandis, 2016). Theoretically, the lifetime of SO$_2$ determined by heterogeneous reaction is negatively correlated with $\gamma$ and $S_p$, and $S_p$ is positively associated with the concentration and $S_{BET}$ of dust as, respectively, described by Eqs. (12) and (10). As a result, the dust with greater heterogeneous reactivity, higher atmospheric loading, or larger $S_{BET}$ is prone to causing the shorter lifetime. During nighttime, IMt causes the shortest SO$_2$ lifetime (46.90 d) due to its
highest concentration relative to the other clay minerals. The relatively large heterogeneous uptake capacity of NAu and SWy links to the second- and third-shortest SO2 lifetimes (221.30 and 260.12 d, respectively). On the other hand, the weakest heterogeneous reactivity of KGa leads to the second-longest lifetime (1758.65 d), and the longest result caused by CCa (2256.25 d) can be interpreted by its lowest S_{BET} that causes the lowest S_p. The presence of solar irradiation alters the lifetime ranking, Imt (20.39 d) < SWy (256.96 d) < NAu (437.85 d) < KGa (597.54 d) < CCa (1488.46 d), as influenced by the different photo-activities of the clay minerals. The heterogeneous drivers of the natural surface are comparable with TMI-O2 in the loss of SO2, and both of them function as the most important lifespan influencers in the absence of sunlight (Fig. S9a). During daytime, the natural dust-driven pathway is only next to the oxidations induced by OH, TMI-O2, and H2O2 (Fig. S9b). The comparison results of the lifetime agree well with those of the sulfate formation rate (Fig. 6), illustrating that the heterogeneous drivers of the dust surface are somewhat responsible for altering the concentrations of gas-phase and particulate sulfur species.

Gas- and aqueous-phase pathways serve as the most significant influencers of the diurnal lifespan of SO2, followed by the dust-driven heterogeneous pathway, while during nighttime these three present closer impacts (Fig. S10). The atmospheric lifetimes of SO2 induced by the considered atmospheric oxidation pathways are calculated as 6.17 d during nighttime and 0.99 d during daytime. Neglecting dust heterogeneous pathways may lengthen the SO2 lifespans to 10.27 and 1.12 d in the absence and presence of solar irradiation, respectively. Analogously, scientists obtained the shorter lifetime of SO2 from the climate model after considering the heterogeneous reaction occurring on volcanic ash particles (Zhu et al., 2020). Clay minerals are more concentrated in the troposphere than volcanic ash and thus have more significant impacts on the removal of atmospheric SO2.

3.4 Sensitivity analysis

The aforementioned calculations set the concentration of natural dust to 55 µg m−3. In contrast to the common atmospheric loading, the burst of sandstorm was normally accompanied by the quickly elevated dust concentration of up to thousands of micrograms per cubic metre (G. Li et al., 2021; Yin et al., 2021; Filonchyk, 2022). It would be meaningful to estimate the heterogeneous contributions in the dust-rich environments. Theoretically, dust concentrations within the ranges of 72–770 µg m−3 during nighttime and 24–260 µg m−3 during daytime could cause the extra sulfate formation of 0.3–3.0 µg m−3 h−1 (Fig. 9a), in line with the acknowledged range of the missing sulfate formation rate (Cheng et al., 2016; T. Liu et al., 2020), which was found to be positively correlated with PM concentration as well (Cheng et al., 2016). Therefore, the heterogeneous reaction of SO2 on the dust surface is a considerable sulfate formation pathway and may evolve into the missing sulfate source in the atmosphere.

The occurrence of sandstorms, particularly during nighttime, aggravates the sulfate pollution in the coarse aerosol mode (Fig. 9b). For instance, the dust concentration of 200 µg m−3, which approaches the PM10 level in North China in March 2021 (Yin et al., 2021), could heterogeneously explain 44.9% of the secondary sulfate during nighttime as well as 29.6% during daytime. It is worthwhile noting that the heterogeneous contribution proportion is susceptible to the evolution of a relatively low dust concentration, and a further increase in dust loading will not significantly elevate the heterogeneous proportion, resulting in a plateau (nighttime) or decrease (daytime) under severe dust pollution. In fact, the increased dust concentration facilitates the aqueous-phase TMI-O2 pathway and dust-mediated and dust-driven heterogeneous reactions but presents negative impacts on the others by the removal of gaseous oxidants over the dust surface. The unsusceptible response to dust concentration, as shown by Fig. 9b, is related to the increased sulfate contributions from TMI-O2. Unlike the TMI-derived oxidation and dust-driven pathway receiving constant positive feedbacks from the increased dust concentration, the dust-mediated pathway is somewhat affected by the decreased oxidants adsorbed on the dust surface. Since the importance of the dust-mediated pathway is more significant during daytime than nighttime, there is a negative correlation between the high dust concentration and heterogeneous contribution proportion in the presence of solar irradiation.

The increased dust concentration, in fact, not only facilitates the heterogeneous process by providing more reactive surfaces, but also affects the gas- and aqueous-phase reactions by altering the atmospheric abundances of reactive species, as explained below. The evolution of dust pollution from slight to heavy conditions would cause the loss of various gaseous oxidants by heterogeneous uptake, and therefore the gas- and aqueous-phase sulfate fluxes, except those induced by TMI-O2, decrease against dust concentration (Bian and Zender, 2003; Tang et al., 2017). Furthermore, the dissolution of mineral constituents produces TMIs in aerosol liquid media (Alexander et al., 2009; Shao et al., 2019), and the irradiated mineral dust was reported to emit gaseous ROS by surface photocatalysis (Dupart et al., 2012; Chen et al., 2021). Herein, the studied oxidation pathways are distinguished by their contribution proportions (Fig. 9c and d). The contribution proportions of OH and H2O2 decrease against dust loading, whereas the importance of TMI-derived oxidation and the dust-driven pathway become more significant as the dust concentrates. The contribution proportion assigned to the dust-mediated process increases with dust concentration at first and then decreases. While the increased dust loading provides more physical space for the dust-mediated reactions, the simultaneously decreased gas-phase oxidants restrain the accumulation of particle-phase oxidants.

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Figure 6. Contributions of diverse atmospheric oxidation pathways to secondary sulfate aerosols. Gas-phase oxidation can be induced by the hydroxyl radical (OH), stabilized Criegee intermediates (CIs), as well as nitrate radical (NO$_3$) only for nighttime. Aqueous-phase oxidation can be induced by hydrogen peroxide (H$_2$O$_2$), NO$_2$, O$_3$, HOCI, HOBr, CH$_3$OOH, CH$_3$COOOH, HONO, TMI-O$_2$, and the T$^*$ and P$_{NO_3}^*$ only for daytime. Dust-mediated heterogeneous oxidation can be initiated by the surface oxidants (H$_2$O$_2$, NO$_2$, O$_3$, HOCl, HOBr, CH$_3$OOH, CH$_3$COOOH, HONO) co-adsorbed with SO$_2$. Dust-driven heterogeneous oxidation can be ascribed to the heterogeneous drivers (transition-metal-bearing components and water-soluble ions) on the surfaces of natural dust and clay minerals – nontronite (NAu), chlorite (CCa), montmorillonite (SWy), kaolin (KGa), and illite (IMt). The (a–d) nighttime and (e–h) daytime conditions were distinguished by the different parameterizations. (a, e) Particle acidity-dependent sulfate formation rates of the diverse atmospheric oxidation pathways. The elements’ colours and shapes are characterized by the legends in solid boxes. (b, f) Quantified sulfate contribution proportions of the studied gas-phase (red), aqueous-phase (blue), and heterogeneous (yellow) reaction pathways. Sulfate formation rates of the dust-mediated and dust-driven pathways during (c–d) nighttime and (g–h) daytime. The effects of ionic strength on the aqueous-phase oxidation were not taken into account. The dust concentration was set to 55 µg m$^{-3}$, representative of the common atmospheric condition of North China (X. Y. Zhang et al., 2012). The panels in dashed boxes right of the legends illustrate the primary physical–chemical processes of atmospheric sulfate formation. More parameterization and methodology details can be found in Texts S2–S4 of the Supplement and Sect. 2.1–2.5 of the main content.

The atmospheric lifetime of SO$_2$ is also affected by the concentration of dust. As shown by Fig. 10a, the lifespan of airborne SO$_2$ during nighttime is higher than that during daytime, and both of them decrease against dust concentration. Analogous to the heterogeneous contribution proportion (Fig. 9b), the lifespan of SO$_2$ is more susceptible to the variation of dust concentration in clean and slightly polluted environments than that under heavily polluted conditions. The mild dust pollution, especially its level variations, should be given more attention. The heterogeneous loss of SO$_2$ by the dust surface was normally evaluated against the gas-phase loss by OH (Ullerstam et al., 2003; Adams et al., 2005; L. Li et al., 2006; Huang et al., 2015; Ma et al., 2018). Current estimation indicates that the airborne dust with the concentration of 45 µg m$^{-3}$ during nighttime or 91 µg m$^{-3}$ during daytime can be regarded as comparable with OH in controlling the removal of SO$_2$ (Fig. 10b). Such dust concentrations are sometimes common in the troposphere, especially during the dust storm periods (G. Li et al., 2021; Yin et al., 2021; Filonchyk, 2022) or near the dust source regions (Ke et al., 2022). Therefore, the heterogeneous loss of SO$_2$ by the airborne dust surface may have a similar magnitude to the main gas-phase loss process and can be taken as an important sink for SO$_2$. 

3.5 Uncertainty analysis

The contribution proportion of dust heterogeneous oxidation could be overestimated or underestimated when considering the uncertainty factors. Herein, the joint impacts of ionic strength (I) and aerosol liquid water content (ALWC) on the aqueous-phase oxidation and the kinetic comparison between microdroplet interfacial oxidation and dust-driven heterogeneous oxidation are discussed to further understand the atmospheric relevance of the heterogeneous drivers for the dust surface under the complex atmospheric conditions.

The aqueous-phase oxidation of SO\(_2\) by H\(_2\)O\(_2\), O\(_3\), NO\(_2\), and TMI-O\(_2\) was quantified under different I–ALWC settings. First, the sulfate formation rate was calculated as a function of ionic strength under the studied ALWC of 300 µg m\(^{-3}\) (Fig. S11). TMI-O\(_2\) and H\(_2\)O\(_2\) dominate the liquid oxidation, while the impacts of NO\(_2\) and O\(_3\) only slightly peak at ~1.0 M during nighttime. Specifically, during nighttime, TMI-O\(_2\) dominates the sulfate formation under the relatively low ionic strength, while the contribution of H\(_2\)O\(_2\) exceeds that of TMI-O\(_2\) over the ionic strength of 0.028 M. During daytime, H\(_2\)O\(_2\) is the predominant oxidant within the studied ionic strength range. Relative to the ionic strength-free calculations, the aqueous oxidation could be weakened by the ionic strength lower than 17.8 M during nighttime or 14.3 M during daytime. Such values can be taken as criteria to distinguish the overestimation or underestimation of liquid kinetics under the ALWC of 300 µg cm\(^{-3}\), which was widely used to characterize the haze events in North China (Cheng et al., 2016).

The joint influences of I and ALWC were further considered (Fig. 11). At each ionic strength, sulfate formation rate associates positively with ALWC. At each ALWC, the increase in ionic strength hinders the aqueous oxidation at first and then facilitates this process, as a consequence resulting in
Figure 9. Sensitivity tests of the sulfate formation rate to dust concentration. (a) Sulfate formation rate of diverse dust-related heterogeneous pathways as a function of dust concentration. The grey area suggests that the missing sulfate formation rate ranges from 0.3 to 3 µg m\(^{-3}\) h\(^{-1}\) as a reference (Cheng et al., 2016; T. Liu et al., 2021). (b) Heterogeneous sulfate proportion varying with dust concentration. Secondary sulfate contributions attributed to the studied oxidation pathways during (c) nighttime and (d) daytime.

a threshold line distinguishing the negative or positive effects of ionic strength, as presented by Fig. 11a and b. Furthermore, the I–ALWC relationships of California, USA (Stelson and Seinfeld, 1981), Beijing, China (Song et al., 2021), Mexico City, Mexico (Volkamer et al., 2007; Hennigan et al., 2015), and the nine cities of Germany (Scheinhardt et al., 2013) were found to be located left of the thresholds, indicating the negative effect of ionic strength under the investigated scenarios (Fig. 11c). Calculated by the reported I–ALWC relationships, the dust-mediated pathway contributes 4.3 %–20.1 % of the secondary sulfate during nighttime and 6.8 %–22.0 % during daytime, and the dust-driven pathway accounts for, respectively, 29.1 %–41.6 % and 9.9 %–12.4 % of the sulfate formation in the absence and presence of sunlight (Fig. 11d). Therefore, the heterogeneous contribution proportions in the complex atmospheric environments can be even higher than those estimated by this study.

Besides the humidified gas–solid interface, the gas–liquid interface of microdroplets is another type of medium that supports the heterogeneous formation of sulfate. The oxidation of SO\(_2\) was found to proceed at the interfacial layer of a droplet with higher kinetics than the bulk process (Jayne...
and Davidovits, 1990). Recently, the interfacial roles of O$_2$ (Hung and Hoffmann, 2015; Hung et al., 2018; Chen et al., 2022), NO$_2$ (Liu and Abbatt, 2021; Yu, 2021), and Mn$^{2+}$ (H. Zhang et al., 2021; W. Wang et al., 2021; T. Wang et al., 2022) have been quantitatively described. Figure 12 compares the dust-driven heterogeneous pathway with the documented gas–liquid interfacial oxidations. As reported by Hung et al. (2015, 2018), the acidic droplet interface favours the non-catalysed oxidation by the presence of sufficient SO$_4^{−}$ and SO$_3^{−}$ when pH < 4.0, which can be further explained by the structure differences of water at the interface and in the bulk phase. Recently, Chen et al. (2022) reevaluated the sulfate formation by interfacial O$_2$ within the pH range of 3.5–4.5 and discovered the positive dependence on ionic strength. The oxidation of SO$_2$ by interfacial NO$_2$ displays a similar pH dependence to that which occurs in bulk solution. The secondary sulfate contributed by interfacial Mn$^{2+}$ is associated positively with particle acidity. Generally, the O$_2$ at the acidic interface dominates the oxidation when pH < 4.0, whereas at pH > 4.0, the interfacial oxidation is primarily controlled by Mn$^{2+}$. Comparing the sulfate formation rates under the same aerosol acidity, the dust surface drivers generally present greater reactivity than the interfacial NO$_2$ and O$_2$. Under the ionic strength of 40 M, the reactivity of interfacial O$_2$ can be regarded as equivalent to that of the dust surface drivers during nighttime. Interfacial Mn$^{2+}$ exceeds the dust surface drivers through the kinetic regime. For instance, the Mn$^{2+}$-catalysed oxidation characterized by W. Wang et al. (2021) is 6.5 or 1.7 times more efficient than the dust-driven pathway in sulfate formation during nighttime or daytime, respectively.

The dust-driven heterogeneous pathway here was investigated by the infrared technique focusing on the bulk particle sample rather than air-suspended particles. Up to now, the microscale effects of dust particles have not been systematically studied. Suspended ATD was the subject of a smog chamber study, and its heterogeneous reactivity toward SO$_2$ was characterized by the net uptake coefficient of $1.71 \times 10^{-6}$ under dark conditions (Park and Jang, 2016). Such a kinetic constant is approximately 1 order of magnitude higher than those obtained from the film-based flow tube tests ($\sim 1.75 \times 10^{-7}$) under the parallel experimental conditions (Y. Zhang et al., 2019). Therefore, in the atmosphere, dust surface drivers may kinetically approach the microdroplet interface in the removal of SO$_2$ and the formation of sulfate.

### 4 Conclusions and implications

This study attempted to deeply understand the importance of heterogeneous oxidation, particularly that induced by dust surface drivers, in the loss of airborne SO$_2$ and formation of sulfate aerosols. Laboratory research was performed first to investigate the heterogeneous driving factors and driving forces. Based on the correlation and regression analysis, transition metal elements, particularly Fe for dark conditions and Al and Ti for photo-reaction, were determined to dominate the heterogeneous oxidation, while water-soluble ions present minor influences. A series of empirical equations was developed to kinetically predict the dust-driven process. The Al-, Fe-, and Ti-bearing mineralogical components, as well as their mixtures, are thus recommended as appropriate proxies for the following laboratory research. The dust-driven heterogeneous $γ$ values were calculated as $6.08 \times 10^{-6}$ for dark reactions and $1.14 \times 10^{-5}$ for photo-reactions, corresponding, respectively, to the atmospheric sulfate forma-tion.
Figure 11. Joint influences of ionic strength (I) and aerosol liquid water content (ALWC) on the aqueous-phase oxidation of SO$_2$. Aqueous-phase sulfate formation rate varying with I and ALWC during (a) nighttime and (b) daytime. The dash-dotted lines indicate the thresholds distinguishing the negative and positive effects of ionic strength. (c) Reported I–ALWC relationships versus the nocturnal and diurnal thresholds. The field-observed data were collected from the measurements performed in California, USA (Stelson and Seinfeld, 1981), Beijing, China (Song et al., 2021), Mexico City, Mexico (Volkamer et al., 2007; Hennigan et al., 2015), and the nine cities of Germany (Scheinhardt et al., 2013). (d) Heterogeneous contribution proportions calculated by the reported I–ALWC relationships and parameterization of this study (ionic strength-free settings and an ALWC of 300 µg m$^{-3}$ for aqueous-phase oxidation; dust concentration of 55 µg m$^{-3}$ for heterogeneous oxidation).

A comparison model was further developed to reveal the atmospheric relevance of the heterogeneous drivers for the dust surface. Dust heterogeneous oxidation is suggested to explain 28.6% of the secondary sulfate aerosols during nighttime and 13.1% during daytime, and the dust surface drivers act as the dominant contributors. Furthermore, dust heterogeneous oxidation affects the atmospheric lifetime of SO$_2$. The increased dust concentration may aggravate the secondary sulfate pollution and has the potential to explain the acknowledged missing sulfate formation rate (0.3–3.0 µg m$^{-3}$ h$^{-1}$). The joint effects of I and ALWC on liquid kinetics influence the importance of the heterogeneous pathway. The heterogeneous contribution proportions estimated by the reported I–ALWC relationships are generally greater than those calculated by the parameterization in the current comparison model. Additionally, a dust heterogeneous pathway in the atmosphere is believed to be kinetically comparable with droplet interfacial oxidation.

Overall, this study suggests that the implementation of heterogeneous processes in atmospheric models will vastly im-
prove the agreement between the modelled and observed sulfate concentrations. The dust heterogeneous pathway should be treated as a significant contributor of secondary sulfate. More necessarily, dust surface drivers are needed to be viewed as an important research focus. Zheng et al. (2015) revised the CMAQ model by adding a heterogeneous mechanism and observed the accurate simulation run upon determining the uptake coefficient to be $2.0 \times 10^{-5}$, which is higher than the $\gamma$ values of the dust heterogeneous pathway herein ($7.04 \times 10^{-6}$ during nighttime and $1.55 \times 10^{-5}$ during daytime). Hence, other solid aerosols, such as sea salts (Laskin et al., 2003; Rossi, 2003) and carbonaceous particles (He and He, 2020; F. Zhang et al., 2020), would better be taken into the following heterogeneous discussions together with the dust-related processes we studied.

Meteorological factors impact the atmospheric relevance of the dust heterogeneous pathway. Taking humidity as an example, on the dust surface, the heterogeneous reaction of SO$_2$ is humidity-dependent, and the exact dependence varies with the type of dust and the condition of the reaction (Huang et al., 2015; Park et al., 2017; Urupina et al., 2022). The uptake of gas-phase oxidants over the dust surface is also influenced by humidity (Kumar et al., 2014). For aerosol droplets, increased humidity elevates its liquid volume and radius (Wu et al., 2018; Ding et al., 2019). Aerosol liquid water serves as an efficient medium for multi-phase reactions (Wu et al., 2018; Yue et al., 2019), whereas the radius is negatively associated with the sulfate formation at the droplet interface (Hung et al., 2018; W. Wang et al., 2021; Chen et al., 2022). Furthermore, droplet acidity decreases as the liquid volume increases, thereby increasing the ionization of dissolved sulfur species, followed by the increased sulfate formation rate (Yue et al., 2019; Jin et al., 2020; Gao et al., 2022). Overall, how meteorological factors influence the relative importance of gas-phase, aqueous-phase, and heterogeneous pathways warrants further research.

Heterogeneous laboratory results were scarcely discussed together with other SO$_2$ conversion routes. This study attempted to set an example for kinetically comparing the heterogeneous oxidation characterized by laboratory work with the documented gas- and aqueous-phase data. Relative to the three-dimensional numerical models, the developed comparison model involves more oxidation pathways. Since the aqueous-phase oxidation here is relevant to aerosol droplets rather than cloud/fog droplets, the current parameterization could be more appropriate for simulating the fine particulate matter that was frequently collected by atmospheric observations and compared with modelling data. Furthermore, the
heterogeneous oxidation was classified into dust-mediated and dust-driven modes to better distinguish the key surface impactors. Therefore, this comparison model has advantages over the traditional atmospheric chemistry models and is recommended for the following heterogeneous laboratory research to systematically compare the experimental data with the acknowledged gas-phase, aqueous-phase, and heterogeneous oxidation pathways.

This work broadens the application of the infrared technique in atmospheric laboratory research. Apart from the uptake coefficient calculations normally concerned, this study moved forward to bridge the relationship between particle acidity and sulfate formation rate by analysing the shape and intensity of the infrared spectrum and further compared heterogeneous oxidation with other atmospheric pathways. This research not only provides a promising methodology for the future heterogeneous research utilizing the classic in situ DRIFTS approach, but also helps to take good advantage of the infrared technique in the laboratory studies in relation to atmospheric heterogeneous oxidation.

Data availability. A dataset for this paper can be accessed at: https://data.mendeley.com/datasets/hyvdz7khs6/1 (Wang, 2022).

Supplement. The supplement related to this article is available online at: https://doi.org/10.5194/acp-22-13467-2022-supplement.

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