

Supplement of Atmos. Chem. Phys., 18, 11125–11133, 2018
<https://doi.org/10.5194/acp-18-11125-2018-supplement>
© Author(s) 2018. This work is distributed under
the Creative Commons Attribution 4.0 License.



Supplement of

Technical note: Comparison and interconversion of pH based on different standard states for aerosol acidity characterization

Shiguo Jia et al.

Correspondence to: Xuemei Wang (eciwxm@jnu.edu.cn) and Liming Yang (cheylm@nus.edu.sg)

The copyright of individual parts of the supplement might differ from the CC BY 4.0 License.

Table S1. Summary of reported aerosol pH calculated by thermodynamic models

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
Molality based pH										
1	Beijing, China	2014	PM2.5	E-AIM-IV	F&R	S	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ , K ⁺ (K ⁺ taken as Na ⁺)	NH ₃ , HNO ₃	3.5 to 5.3 for forward mode and -2 to 10 for reverse mode.	(Song et al., 2018)
				ISORROPIA-II	F&R	S&MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻			
2	Guangzhou, China	2013	PM2.5	E-AIM-III	F	S	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl	2.4 ± 0.3	(Jia et al., 2018)
3	Southeastern US	2013	PM1 and PM2.5	AIOMFAC	R	MS	NH ₄ ⁺ , SO ₄ ²⁻ , organic compounds ^c	- g	1.4±1.2 (EQLB) 1.3±1.2 (CLLPS)	(Pye et al., 2018)
				AIOMFAC	R	MS	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ , organic compounds ^c	- g	1.5±2.0 (EQLB) 1.3±2.1 (CLLPS)	
				ISORROPIA-II	R	- g	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- g	0.7±2.5	
				ISORROPIA-II	F	- g	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	1.1±0.7	

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
4	Beijing, China	2015& 2016	PM2.5	ISORROPIA	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ ^d	NH ₃ , NO ₃ , HCl	4.2 ± 0.5	(Liu et al., 2017)
5	Baltimore & Chicago, US	2011 to 2015	PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	~1.0 to 2.7 ^e	(Battaglia et al., 2017)
6	Southeastern US	2010	PM1 & PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl	1.9±0.5 for PM1 and 2.7±0.3 for PM2.5	(Guo et al., 2017a)
7	Beijing & Xi'an, China	2013	PM1 & PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	4.5 for Beijing and 5 for Xi'an	(Guo et al., 2017b)
8	Atlanta, US	2016	<18 μm ^f	ISORROPIA-II	F	- ^g	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl ^h	1 to 2 for fine particle; ~6 to ~7 for coarse particle ^e	(Fang et al., 2017)
9	Finokalia, Greece	2012	PM1 and PM10	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ ^d	NH ₃ ⁱ	0.5 to 2.8	(Bougiatioti et al., 2016)
10	Northeastern US	2015	PM1	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ ^h	0.77 ± 0.96	(Guo et al., 2016)
11	Southeastern US	2012& 2013	PM1 and PM 2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ ^d	NH ₃	0.94 ±0.59	(Guo et al., 2015)
12	Southeastern US	2013	PM1	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	0-2	(Xu et al., 2015)

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
Molarity based pH										
13	Beijing, China	2013	PM2.1	E-AIM-IV	R	- g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- g	1.1	(Tian et al., 2018)
14	Chengdu, China	2011	0.18 to 1.8 μm	E-AIM-II	R	- g	SO ₄ ²⁻ , NH ₄ ⁺ , NO ₃ ⁻	- g	-1.11	(Cheng et al., 2015)
15	Singapore	2011	PM2.5	E-AIM-III	R	- g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- g	0.48 & 0.72 (day & night)	(Behera et al., 2013)
16	Jinan, China	2006 & 2007	PM1	E-AIM-II	R	- g	NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻	- g	-1	(Cheng et al., 2011)
17	Hong Kong	2008 &2009	PM2.5	E-AIM-III	R	- g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- g	-0.45 to 0.59	(Xue et al., 2011)
18	Beijing, China	2005	PM2.5	AIM-II	R	- g	NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻	- g	-0.52±0.62	(Pathak et al., 2009)
	Shanghai, China	2005			R				-0.77±0.67	
	Lanzhou, China	2006,			R				-0.38±0.64	
	Guangzhou, China	2004			R				0.61±0.71	
19	Hong Kong	2001	PM10	AIM-II	R	- g	NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻	- g	0.25	(Pathak et al., 2004)

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
Male fraction based pH										
20	Southeastern US	1999 to 2014	PM2.5	ISORROPIA-II	F	- ^g	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ ⁱ	0-2	(Weber et al., 2016)
21	Mexico City, US	2006	PM2.5	ISORROPIA-II E-AIM-II and IV	R	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , and Cl ⁻ for AIM-IV and NH ₄ ⁺ , SO ₄ ²⁻ and NO ₃ ⁻ for AIM-II	- ^g	1.98	(Hennigan et al., 2015)
				ISORROPIA-II, E-AIM-II and IV	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻ for AIM-IV and NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ for AIM-II	NH ₃ , HNO ₃	3.31	
						S		NH ₃ and HNO ₃	3.24	
22	Po Valley, Italy	2009	PM2.5	E-AIM-IV	R	- ^g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- ^g	2.7 to 3.8	(Squizzato et al., 2013)
23	Beijing, China	2005& 2006	PM2.5	E-AIM-II	R	- ^g	NH ₄ ⁺ , SO ₄ ²⁻ and NO ₃ ⁻	- ^g	1.2 ± 1.1	(He et al., 2012)
24	Pittsburgh, US	2002	PM1	E-AIM-II	R	- ^g	SO ₄ ²⁻ , NH ₄ ⁺ , NO ₃ ⁻	- ^g	~0.5 to 5 ^f	(Zhang et al., 2007)

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
pH not specified										
25	Beijing, China	2014& 2015	PM2.5	ISORROPIA-II	F	S	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ ^j	7.6 ± 0.1	(He et al., 2018)
26	Beijing , China	2014	PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	- g	4.1	(Tan et al., 2018)
27	Tianjin , China	2014& 2015	PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	4.9±0.4	(Shi et al., 2017)
28	Beijing, China	2015	PM1	ISORROPIA-II	F	S ^k	NH ₃ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃	7.6 ± 0.0	(Wang et al., 2016)
	Xi'an, China	2012	PM2.5	ISORROPIA-II	F	S ^k	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻		7.0 ± 1.3	
29	Beijing, China	2013	PM2.5	ISORROPIA-II	F	MS	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ^{-e}	NH ₃ ^j	4.4 ± 0.6	
30	São Paulo, Brazil	2012	<18 μm ^f	E-AIM	R	R	- g	- g	5.4	(Cheng et al., 2016)
								- g	6.2	
31	Hong Kong	1996 to 1998	PM1	AIM-II	R	- g	NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻	- g	4.1 to 5.4 ^c	(Vieira-Filho et al., 2016)
								- g	-2.5 to 1.5	(Yao et al., 2007)

#	Location	Year	Particle size	Model	F/ R ^a	S/ MS ^b	Particulate species	Gases	pH	References
32	Hong Kong	2000	PM2.5	AIM-II	R	- ^g	NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻	- ^g	-0.4	(Yao et al., 2006)
				SCAPE2	F	- ^g	Na ⁺ , K ⁺ , Ca ²⁺ , Mg ²⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl	2.3	
				ISORROPIA-I	F	- ^g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl	3.9	
				SCAPE2	F	- ^g	Na ⁺ , NH ₄ ⁺ , SO ₄ ²⁻ , NO ₃ ⁻ , Cl ⁻	NH ₃ , NO ₃ , HCl	1.9	

Note: ^a F means forward and R means reverse. ^b S means stable and MS means metastable. ^c Organic compounds were considered either in complete liquid–liquid phase separation (CLLPS) or equilibrium (EQLB) mode; Ca²⁺, K⁺ and Mg²⁺ were taken as Na⁺. ^d Aerosol water content associated with organic compounds considered; ^e Data extracted from figure; ^f Size segregated samples; ^g - means not specified; ^h Assuming NH₄⁺ (total) = NH₃ (gas phase) +NH₄⁺ (particle phase); ⁱ Estimated based on iteration; ^j NH₃ estimated based on empirical equation; ^k Obtained from Song et al. (2018).

Table S2. A summary of estimation methods of parameters for pH calculation based on different standard states.

Parameter	E-AIM-IV	ISORROPIA-II	AIOMFAC
Molar fraction			
x_H	output	Eq. (7)	output
f_H	output	1 ^a	Eq. (4)
Molality			
m_H	output	output	output
γ_H	Eq. (4)	1 ^a	output
Molarity			
c_H	Eq. (8) ^b	Eq. (8) ^b	Eq. (8) ^b
y_H	Eq. (6)	1 ^a	Eq. (6)

Note: ^aActivity coefficient is assumed to be 1; ^bthe density of aerosol solution is based on the result from E-AIM-IV.

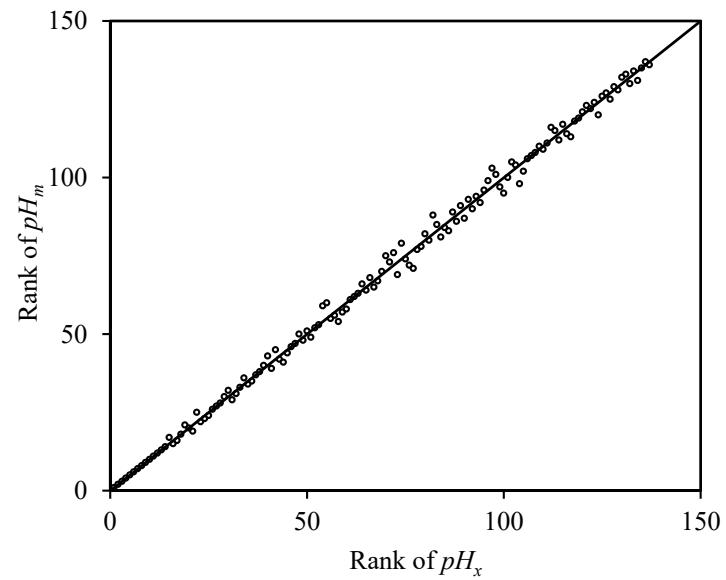


Figure S1. Comparison of the rank of pH_x and pH_m

Reference

- Battaglia, M. A., Douglas, S., and Hennigan, C. J.: Effect of the Urban Heat Island on Aerosol pH, *Environ. Sci. Technol.*, 51, 13095-13103, 10.1021/acs.est.7b02786, 2017.
- Behera, S. N., Betha, R., Liu, P., and Balasubramanian, R.: A study of diurnal variations of PM_{2.5} acidity and related chemical species using a new thermodynamic equilibrium model, *Sci. Total Environ.*, 452–453, 286-295, <http://dx.doi.org/10.1016/j.scitotenv.2013.02.062>, 2013.
- Bougiatioti, A., Nikolaou, P., Stavroulas, I., Kouvarakis, G., Weber, R., Nenes, A., Kanakidou, M., and Mihalopoulos, N.: Particle water and pH in the eastern Mediterranean: source variability and implications for nutrient availability, *Atmos. Chem. Phys.*, 16, 4579-4591, 10.5194/acp-16-4579-2016, 2016.
- Cheng, C., Wang, G., Meng, J., Wang, Q., Cao, J., Li, J., and Wang, J.: Size-resolved airborne particulate oxalic and related secondary organic aerosol species in the urban atmosphere of Chengdu, China, *Atmos. Res.*, 161-162, 134-142, <https://doi.org/10.1016/j.atmosres.2015.04.010>, 2015.
- Cheng, S.-h., Yang, L.-x., Zhou, X.-h., Xue, L.-k., Gao, X.-m., Zhou, Y., and Wang, W.-x.: Size-fractionated water-soluble ions, situ pH and water content in aerosol on hazy days and the influences on visibility impairment in Jinan, China, *Atmos. Environ.*, 45, 4631-4640, <https://doi.org/10.1016/j.atmosenv.2011.05.057>, 2011.
- Cheng, Y., Zheng, G., Wei, C., Mu, Q., Zheng, B., Wang, Z., Gao, M., Zhang, Q., He, K., Carmichael, G., Pöschl, U., and Su, H.: Reactive nitrogen chemistry in aerosol water as a source of sulfate during haze events in China, *Science Advances*, 2, 10.1126/sciadv.1601530, 2016.
- Fang, T., Guo, H., Zeng, L., Verma, V., Nenes, A., and Weber, R. J.: Highly Acidic Ambient Particles, Soluble Metals, and Oxidative Potential: A Link between Sulfate and Aerosol Toxicity, *Environ. Sci. Technol.*, 51, 2611-2620, 10.1021/acs.est.6b06151, 2017.
- Guo, H., Xu, L., Bougiatioti, A., Cerully, K. M., Capps, S. L., Hite Jr, J. R., Carlton, A. G., Lee, S. H., Bergin, M. H., Ng, N. L., Nenes, A., and Weber, R. J.: Fine-particle water and pH in the southeastern United States, *Atmos. Chem. Phys.*, 15, 5211-5228, 10.5194/acp-15-5211-2015, 2015.
- Guo, H., P., S. A., Pedro, C. J., C., S. J., D., L. H. F., E., D. J., L., J. J., A., T. J., S., B. S., Athanasios, N., and J., W. R.: Fine particle pH and the partitioning of nitric acid during winter in the northeastern United States, *J. Geophys. Res.: Atmos.*, 121, 10,355-310,376, doi:10.1002/2016JD025311, 2016.
- Guo, H., Liu, J., Froyd, K. D., Roberts, J. M., Veres, P. R., Hayes, P. L., Jimenez, J. L., Nenes, A., and Weber, R. J.: Fine particle pH and gas-particle phase partitioning of inorganic species in Pasadena, California, during the 2010 CalNex campaign, *Atmos. Chem. Phys.*, 17, 5703-5719, 10.5194/acp-17-5703-2017, 2017a.
- Guo, H., Weber, R. J., and Nenes, A.: High levels of ammonia do not raise fine particle pH sufficiently to yield nitrogen oxide-dominated sulfate production, *Scientific Reports*, 7, 12109, 10.1038/s41598-017-11704-0, 2017b.
- He, K., Zhao, Q., Ma, Y., Duan, F., Yang, F., Shi, Z., and Chen, G.: Spatial and seasonal variability of PM_{2.5} acidity at two Chinese megacities: insights into the formation of secondary inorganic aerosols, *Atmos. Chem. Phys.*, 12, 1377-1395, 10.5194/acp-12-1377-2012, 2012.
- He, P., Alexander, B., Geng, L., Chi, X., Fan, S., Zhan, H., Kang, H., Zheng, G., Cheng, Y., Su, H., Liu, C., and Xie, Z.: Isotopic constraints on heterogeneous sulfate production in Beijing haze, *Atmos. Chem. Phys.*, 18, 5515-5528, 10.5194/acp-18-5515-2018, 2018.
- Hennigan, C. J., Izumi, J., Sullivan, A. P., Weber, R. J., and Nenes, A.: A critical evaluation of proxy methods used to estimate the acidity of atmospheric particles, *Atmos. Chem. Phys.*, 15, 2775-2790, 10.5194/acp-15-2775-2015, 2015.
- Jia, S., Sarkar, S., Zhang, Q., Wang, X., Wu, L., Chen, W., Huang, M., Zhou, S., Zhang, J., Yuan, L., and Yang, L.: Characterization of diurnal variations of PM_{2.5} acidity using an open thermodynamic system: A case study of Guangzhou, China, *Chemosphere*, 202, 677-685, <https://doi.org/10.1016/j.chemosphere.2018.03.127>, 2018.
- Liu, M., Yu, S., Tian, Z., Zhenying, X., Caiqing, Y., Mei, Z., Zhijun, W., Min, H., Yusheng, W., and Tong, Z.: Fine particle pH during severe haze episodes in northern China, *Geophys. Res. Lett.*, 44, 5213-5221, doi:10.1002/2017GL073210, 2017.
- Pathak, R. K., Yao, X., and Chan, C. K.: Sampling Artifacts of Acidity and Ionic Species in PM_{2.5}, *Environ. Sci. Technol.*, 38, 254-259, 10.1021/es0342244, 2004.
- Pathak, R. K., Wu, W. S., and Wang, T.: Summertime PM_{2.5} ionic species in four major cities of China: nitrate formation in an ammonia-deficient atmosphere, *Atmos. Chem. Phys.*, 9, 1711-1722, 10.5194/acp-9-1711-2009, 2009.

- Pye, H. O. T., Zuend, A., Fry, J. L., Isaacman-VanWertz, G., Capps, S. L., Appel, K. W., Foroutan, H., Xu, L., Ng, N. L., and Goldstein, A. H.: Coupling of organic and inorganic aerosol systems and the effect on gas-particle partitioning in the southeastern US, *Atmos. Chem. Phys.*, 18, 357-370, 10.5194/acp-18-357-2018, 2018.
- Shi, G., Xu, J., Peng, X., Xiao, Z., Chen, K., Tian, Y., Guan, X., Feng, Y., Yu, H., Nenes, A., and Russell, A. G.: pH of Aerosols in a Polluted Atmosphere: Source Contributions to Highly Acidic Aerosol, *Environ. Sci. Technol.*, 51, 4289-4296, 10.1021/acs.est.6b05736, 2017.
- Song, S., Gao, M., Xu, W., Shao, J., Shi, G., Wang, S., Wang, Y., Sun, Y., and McElroy, M. B.: Fine-particle pH for Beijing winter haze as inferred from different thermodynamic equilibrium models, *Atmos. Chem. Phys.*, 18, 7423-7438, 10.5194/acp-18-7423-2018, 2018.
- Squizzato, S., Masiol, M., Brunelli, A., Pistollato, S., Tarabotti, E., Rampazzo, G., and Pavoni, B.: Factors determining the formation of secondary inorganic aerosol: a case study in the Po Valley (Italy), *Atmos. Chem. Phys.*, 13, 1927-1939, 10.5194/acp-13-1927-2013, 2013.
- Tan, T., Hu, M., Li, M., Guo, Q., Wu, Y., Fang, X., Gu, F., Wang, Y., and Wu, Z.: New insight into PM2.5 pollution patterns in Beijing based on one-year measurement of chemical compositions, *Sci. Total Environ.*, 621, 734-743, <https://doi.org/10.1016/j.scitotenv.2017.11.208>, 2018.
- Tian, S., Pan, Y., and Wang, Y.: Ion balance and acidity of size-segregated particles during haze episodes in urban Beijing, *Atmos. Res.*, 201, 159-167, <https://doi.org/10.1016/j.atmosres.2017.10.016>, 2018.
- Vieira-Filho, M., Pedrotti, J. J., and Fornaro, A.: Water-soluble ions species of size-resolved aerosols: Implications for the atmospheric acidity in São Paulo megacity, Brazil, *Atmos. Res.*, 181, 281-287, <https://doi.org/10.1016/j.atmosres.2016.07.006>, 2016.
- Weber, R. J., Guo, H., Russell, A. G., and Nenes, A.: High aerosol acidity despite declining atmospheric sulfate concentrations over the past 15 years, *Nature Geosci.*, 9, 282-285, 10.1038/ngeo2665, 2016.
- Xu, L., Guo, H., Boyd, C. M., Klein, M., Bougiatioti, A., Cerully, K. M., Hite, J. R., Isaacman-VanWertz, G., Kreisberg, N. M., Knote, C., Olson, K., Koss, A., Goldstein, A. H., Hering, S. V., de Gouw, J., Baumann, K., Lee, S.-H., Nenes, A., Weber, R. J., and Ng, N. L.: Effects of anthropogenic emissions on aerosol formation from isoprene and monoterpenes in the southeastern United States, *Proceedings of the National Academy of Sciences*, 112, 37-42, 10.1073/pnas.1417609112, 2015.
- Xue, J., Lau, A. K. H., and Yu, J. Z.: A study of acidity on PM2.5 in Hong Kong using online ionic chemical composition measurements, *Atmos. Environ.*, 45, 7081-7088, <j.atmosenv.2011.09.040>, 2011.
- Yao, X., Yan Ling, T., Fang, M., and Chan, C. K.: Comparison of thermodynamic predictions for in situ pH in PM2.5, *Atmos. Environ.*, 40, 2835-2844, <j.atmosenv.2006.01.006>, 2006.
- Yao, X., Ling, T. Y., Fang, M., and Chan, C. K.: Size dependence of in situ pH in submicron atmospheric particles in Hong Kong, *Atmos. Environ.*, 41, 382-393, <j.atmosenv.2006.07.037>, 2007.
- Zhang, Q., Jimenez, J. L., Worsnop, D. R., and Canagaratna, M.: A Case Study of Urban Particle Acidity and Its Influence on Secondary Organic Aerosol, *Environ. Sci. Technol.*, 41, 3213-3219, 10.1021/es061812j, 2007.