



HO₂NO₂ and HNO₃ in the coastal Antarctic winter night: a “lab-in-the-field” experiment

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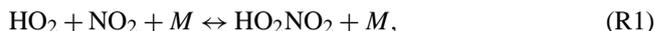
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Abstract. Observations of peroxyntic acid (HO₂NO₂) and nitric acid (HNO₃) were made during a 4 month period of Antarctic winter darkness at the coastal Antarctic research station, Halley. Mixing ratios of HNO₃ ranged from instrumental detection limits to ~8 parts per trillion by volume (pptv), and of HO₂NO₂ from detection limits to ~5 pptv; the average ratio of HNO₃ : HO₂NO₂ was 2.0(±0.6) : 1, with HNO₃ always present at greater mixing ratios than HO₂NO₂ during the winter darkness. An extremely strong association existed for the entire measurement period between mixing ratios of the respective trace gases and temperature: for HO₂NO₂, $R^2 = 0.72$, and for HNO₃, $R^2 = 0.70$. We focus on three cases with considerable variation in temperature, where wind speeds were low and constant, such that, with the lack of photochemistry, changes in mixing ratio were likely to be driven by physical mechanisms alone. We derived enthalpies of adsorption (ΔH_{ads}) for these three cases. The average ΔH_{ads} for HNO₃ was $-42 \pm 2 \text{ kJ mol}^{-1}$ and for HO₂NO₂ was $-56 \pm 1 \text{ kJ mol}^{-1}$; these values are extremely close to those derived in laboratory studies. This exercise demonstrates (i) that adsorption to/desorption from the snow pack should be taken into account when addressing budgets of boundary layer HO₂NO₂ and HNO₃ at any snow-covered site, and (ii) that Antarctic winter can be used as a natural “laboratory in the field” for testing data on physical exchange mechanisms.

1 Introduction

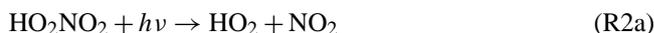
Peroxyntic acid (HO₂NO₂, also written as HNO₄) and nitric acid (HNO₃) are acidic gases that are of increasing interest to polar tropospheric chemistry. Their primary relevance is that they act as reservoir species for HO_x and NO_x, which are now recognised to drive the surprisingly vigorous oxidation chemistry that has been observed during Antarctic summer (e.g. Davis et al., 2001; Chen et al., 2001). The spatial and temporal distribution of HO₂NO₂ and HNO₃ across the polar regions thus becomes important for understanding the overall atmospheric chemical system, and models require details of their sources, and any physical exchange process by which they move from one environmental compartment to another. Currently, many of these details are unknown.

The gas-phase chemistry of HO₂NO₂ and HNO₃ is relatively straightforward. Peroxyntic acid is a somewhat unstable molecule that forms and dissociates through its temperature-dependent equilibrium reaction:



which renders an increased stability for HO₂NO₂ at lower temperatures.

There are a number of photodissociation pathways which drive HO₂NO₂ and HNO₃ chemistry under sunlit conditions. The most important are thought to be



Peroxyntic acid can also be lost through reaction with OH:



Gas-phase production of nitric acid proceeds via



The major loss processes are reaction with OH and photolysis:



Both HO₂NO₂ (Ulrich et al., 2012) and HNO₃ (Bartels-Rausch et al., 2002; Hudson et al., 2002; Ullerstam et al., 2005) have been shown in laboratory experiments to adsorb to ice surfaces. This conclusion is supported by field observations which have confirmed uptake of both HNO₃ and HO₂NO₂ to snow surfaces (Huey et al., 2004; Slusher et al., 2002), and of HNO₃ to cirrus clouds (Weinheimer et al., 1998; Popp et al., 2004; Ziereis et al., 2004). In general, therefore, in snow covered areas, or indeed regions of the atmosphere with lofted snow/ice, such as cirrus clouds or blowing/precipitating snow, physical adsorption of HNO₃ and HO₂NO₂ from the air to the snow/ice is likely to occur. The details of this uptake will differ somewhat between the two molecules because the partitioning coefficient of HNO₃ is greater than that for HO₂NO₂. HNO₃ is thus more sticky than HO₂NO₂, and a higher fraction of HNO₃ can be expected on the ice surface compared with HO₂NO₂.

High resolution observations of HNO₃ and HO₂NO₂ in the polar regions are scarce. Critically, both HO₂NO₂ and HNO₃ have been measured together during a number of Antarctic studies at high temporal resolution. These studies have included both ground-based experiments at the South Pole (Slusher et al., 2002; Huey et al., 2004) and airborne measurements across the wider Antarctic Plateau (Slusher et al., 2010).

The ground-based studies have revealed considerable inter-annual variability in summertime HNO₃ and HO₂NO₂ mixing ratios, but always of the order of 10 s of pptv at the South Pole. For example, the median observed HNO₃ between 16 and 31 December 2000 was 18.2 pptv and for HO₂NO₂ was 23.5 pptv (Davis et al., 2004); over the equivalent time period in 2003, the median HNO₃ was 84 pptv, and for HO₂NO₂ was 39 pptv (Eisele et al., 2008). Considerably greater mixing ratios have also been observed; for example, the median mixing ratio of HNO₃ between 15 and 30 November 2003 was 194 pptv, and of HO₂NO₂ was 63 pptv (Eisele et al., 2008). While the specific mixing ratios will be strongly influenced by boundary layer height, overall, these high mixing ratios are fuelled by in situ production from elevated levels of NO_x and HO_x within the South Pole boundary layer, in turn driven by photochemical release of trace gases from the surrounding snowpack (Davis et al., 2001, 2008).

The airborne measurements assessed the three-dimensional distribution of HO₂NO₂ and HNO₃ across

the Antarctic Plateau during the ANTCI 2005 campaign (Slusher et al., 2010). They revealed significant vertical gradients in both species, with higher concentrations at the ground, consistent with a source associated with emissions from the snowpack. The measurements also showed a widespread distribution of both HNO₃ and HO₂NO₂ across the Plateau region.

To date there have been no measurements of high temporal resolution HO₂NO₂ and HNO₃ in coastal Antarctica, and no measurements at all from Antarctica outside the summer season. We report here observations of HO₂NO₂ and HNO₃ made using a chemical ionisation mass spectrometer (CIMS) at Halley research station in coastal Antarctica (75°35' S, 26°39' W) from 24 May to 18 September 2007. The data allow us to assess whether HNO₃ and HO₂NO₂ are present in significant concentrations at other Antarctic locations and seasons than the Antarctic Plateau in summer. They also provide an opportunity to test laboratory-derived physical exchange parameters under semi-constrained, but genuine real-world conditions. At Halley, the sun remains below the horizon from 30 April to 13 August, such that this new data set includes many weeks of winter darkness. Under these conditions of 24 h per day darkness, atmospheric photochemistry stalls, and trace gas concentrations are controlled entirely by either transport or physical air–snow exchange.

2 Experimental

2.1 CIMS instrumentation

The CIMS instrument used in this study has been described in detail elsewhere (Buys et al., 2013). It was installed in the Clean Air Sector Laboratory (CASLab), which is located roughly 1 km from the main Halley station, and in a sector that rarely receives air from the base (Jones et al., 2008). The CIMS inlet extended ~20 cm above the roof of CASLab, at a height roughly 5 m above the surrounding snowpack. The inlet system was designed to minimise residence time and surface losses (Neuman et al., 1999).

The instrument employed the SF₆⁻ method to detect both HNO₃ and HO₂NO₂, using the NO₄⁻ (HF) cluster at *m/z* 98 to detect HO₂NO₂, and NO₃⁻ (HF) at *m/z* 82 to detect HNO₃ (as per Slusher et al., 2001, 2002). Calibration was achieved using the SO₂ method as described by Slusher et al. (2001) and Kim et al. (2007). Background measurements, or zeros, were obtained every 10 min. These were achieved by passing sampled air for 3 min through a customised filter filled with activated coarse charcoal and nylon glass wool coated in NaHCO₃. This scrubbing method has previously been shown to be efficient at removing both HO₂NO₂ and HNO₃ from sampled air (Slusher et al., 2001). The instrument detection limit derived from background data averaged over 10 min was 0.7 pptv for HNO₃ and 0.4 pptv for HO₂NO₂. Total estimated uncertainty in the CIMS observations is ±40 %.

While the SF₆⁻ method has been used successfully in previous field campaigns (e.g. Slusher et al., 2002, 2010), it has been demonstrated in laboratory studies (Slusher et al., 2001) that SF₆⁻ reacts with H₂O in the sample air flow. This introduces an interferent into the technique, the non-linearity of which is evident in the unfiltered data (not shown). However, with their instrument reaction time of ~25 ms, Slusher et al. (2001) also concluded that this interferent was significant only at dew points greater than -25 °C, and that at lower dew points, the interferent was negligible. During the period of measurements at Halley, the CIMS instrument also operated with a reaction time of ~25 ms, such that the interferent would be equivalent to that of Slusher et al. (2001). At Halley, dewpoint temperatures varied from -12 to -52 °C (mean -31 °C), but were below -25 °C for 81 % of the time. To remove the potential for H₂O interference in our data, all measurements made at dew points above -25 °C are filtered out from the data set.

2.2 Boundary layer meteorology

Measurements of near-surface boundary layer meteorology were made on a 32 m profiling mast located ~25 m from the CASLab. Bulk sensors were located at 1, 2, 4, 8, 16, and 32 m above ground level, recording at 1 Hz and averaged to 10 min means; temperatures and humidity were measured with platinum resistance thermometers (0.1 K resolution) and solid state humidity probes (2 % resolution), respectively, using an aspirated HMP35D from Vaisala Corp. 10 min vector average wind speed and direction were measured with R. M. Young propeller vanes at 0.1 ms⁻¹ and 2° resolution.

Three three-axis Metek USA-1 ultrasonic anemometer/thermometers were deployed at the 4, 16, and 32 m levels, sampling at 20 Hz. The data were tilt corrected and the relevant co-variances calculated over 1 min means.

3 Results and discussion

3.1 Overall data series

The time series of filtered HO₂NO₂ and HNO₃ data, averaged to hourly means, is shown in Fig. 1, together with hourly means of ambient temperature, dew point temperature, and solar zenith angle. The solar zenith angle shows that the sun was below the horizon for the majority of this time period, and observations of NO and NO₂ (made using a Sonoma Tech. Dual channel chemiluminescence analyser (Bauguitte et al., 2012; Cotter et al., 2003), with detection limits of 2 pptv for NO and 6 pptv for NO₂, not shown) were consistently below the instrumental detection limits.

Regardless of this apparent lack of photochemical activity, there is considerable variability in the HNO₃ and HO₂NO₂ observations. Mixing ratios of HNO₃ ranged from instrumental detection limits to ~8 parts per trillion by volume (pptv) and of HO₂NO₂ varied from detection limits to

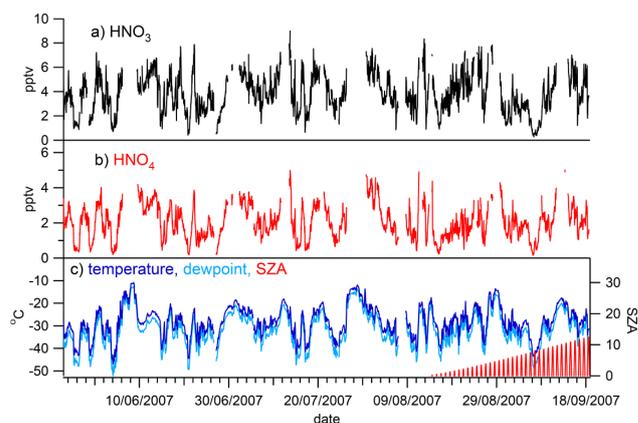


Figure 1. Time series of (a) HNO₃, (b) HO₂NO₂, and (c) ambient temperature, dewpoint, and solar zenith angle (SZAs) (hourly averages) for the entire measurement period discussed in this paper, 24 May to 18 September 2007.

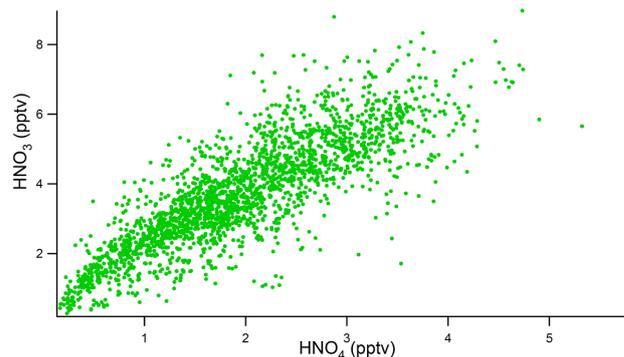


Figure 2. Nitric acid (HNO₃) vs. peroxyntic acid (HO₂NO₂); hourly averages of measurements made from 24 May to 18 September 2007.

~5 pptv. These values are considerably lower than those observed at the South Pole in summer, where photochemical production is fuelled by emissions of NO_x from the snowpack, and where mixing ratios of HNO₃ and HO₂NO₂ were generally in the 10 s of pptv, and sometimes over 100 pptv, as discussed earlier.

It is also noticeable in the Halley data that the pattern of variability in both HO₂NO₂ and HNO₃ was very similar, with each time series tracking the other closely. Indeed the correlation between the two chemical species was high, as shown in Fig. 2, where the correlation coefficient, R^2 , for the hourly averages was 0.70. This finding is consistent with the data of Slusher et al. (2002), which showed that, although the range of mixing ratios at the South Pole in summer were considerably higher than observed at Halley during the winter (<5 to 54 pptv for HO₂NO₂, and <5 to 68 pptv for HNO₃), the variability observed in both species during the measurement period was, again, highly coupled.

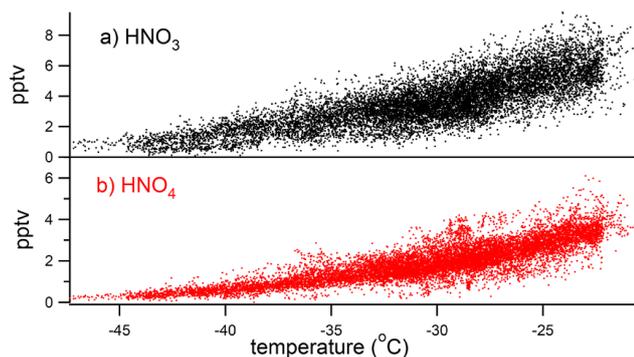


Figure 3. (a) Nitric acid (HNO₃) and (b) peroxyxynitric acid (HO₂NO₂) vs. ambient temperature; 10 min averages of measurements made from 24 May to 18 September 2007.

The time series presented in Fig. 1 also clearly shows the very strong association between mixing ratios of HNO₃ and HO₂NO₂ and ambient (and dewpoint) temperature. This association is further demonstrated in Fig. 3, which shows both HNO₃ and HO₂NO₂ plotted against ambient temperature, again for the entire period of measurements. Calculated correlation coefficients with temperature are high, with $R^2 = 0.70$ and 0.72 for HNO₃ and HO₂NO₂ respectively.

Given that mixing ratios of HNO₃ and HO₂NO₂ are so strongly associated with ambient temperature, Table 1 gives the mean and standard deviation of HNO₃ and HO₂NO₂ measured between 24 May and 18 September 2007, calculated within specific temperature ranges. The statistics are derived using only the filtered data that were above 3σ detection limits. At the South Pole during the summer, ambient temperature ranged from -31.5 to -23.6 °C, with a mean of -27.7 °C; within this temperature range, HNO₃ mixing ratios ranged from < 5 to 54 pptv (mean 25 pptv) and HO₂NO₂ ranged from < 5 to 68 pptv (mean 22 pptv). For the equivalent temperature range at Halley (also with a mean of -27.7 °C), mean and maximum mixing ratios for HNO₃ were 4.4 and 9 pptv respectively, and for HO₂NO₂ were 2.5 and 5 pptv respectively, clearly significantly below those observed at the South Pole, for the reasons outlined above.

Differences between South Pole summer and Halley winter are also evident in the ratio of HNO₃:HO₂NO₂, arising through differences in the species' lifetimes. Throughout the Halley measurement period, the average ratio of HNO₃:HO₂NO₂ was $2.0(\pm 0.6):1$, with HNO₃ always (apart from a few outliers) present at greater mixing ratios than HO₂NO₂ during the winter darkness (see Fig. 4). This finding is in contrast to observations from the South Pole during sunlit summer time, when mixing ratios of HNO₃ and HO₂NO₂ were roughly equal for much of the measurement period (Slusher et al., 2002). Figure 1 of Slusher et al. (2002) shows that HO₂NO₂ was present at higher mixing ratios than HNO₃ for roughly 2 out of the 7 days of measurements; during the roughly 4 months of measurements at Halley, the only

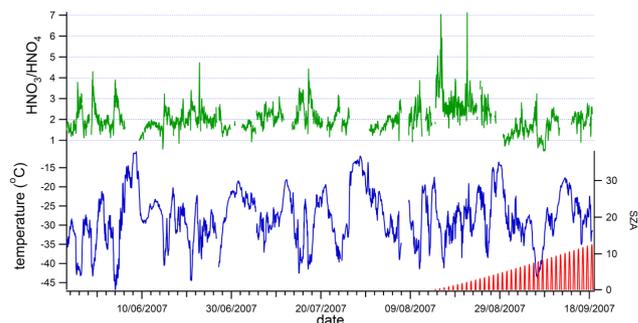


Figure 4. Ratio of HNO₃:HO₂NO₂ (hourly average data) for the May to September measurement period. Also shown, for reference, are ambient temperature and solar zenith angle.

occasion when the mixing ratio of HO₂NO₂ exceeded that of HNO₃ was on the 6 and 7 September, a period when temperatures were particularly low but there was a limited amount of sunlight.

3.2 Short-term variability in HNO₃ and HO₂NO₂ and link to ambient temperature

The short-term variability in the HNO₃ and HO₂NO₂ is shown more clearly in Fig. 5. The three examples show periods when ambient air temperatures varied rapidly and considerably, but where they remained below the -25 °C dewpoint threshold such that no chemical data filtering was required. These 10 min averages show that even very small-scale features of temperature change are reflected in the chemical measurements. For example, at midnight on 5 June, the short-lived peak in temperature is reflected also in HNO₃ and HO₂NO₂; the temperature peak around 11 a.m. on 21 June is apparent with similar, small, peaks in HNO₃ and HO₂NO₂; and the short-lived temperature peak around noon on 15 July is also evident in short-lived increases in HNO₃ and HO₂NO₂ mixing ratios. While large-scale variability in HNO₃ and HO₂NO₂ could be linked to air mass origin, such fine-scale variability can only be explained by a local, fast-acting, source/sink mechanism. The association between variability in HNO₃ and HO₂NO₂ and changes in ambient temperature strongly suggest a temperature-dependent mechanism. Given our understanding of the interaction between acidic gases and ice gained through laboratory studies (e.g. Huthwelker et al., 2006), one possible mechanism is temperature-dependent adsorption/desorption at the snow surface.

3.3 Evidence for HO₂NO₂ and HNO₃ air/snow exchange

To probe in more detail the response of HNO₃ and HO₂NO₂ to changes in temperature, we examined periods in the data where ambient temperatures changed, but where wind speeds were relatively low and invariable. By adopting this ap-

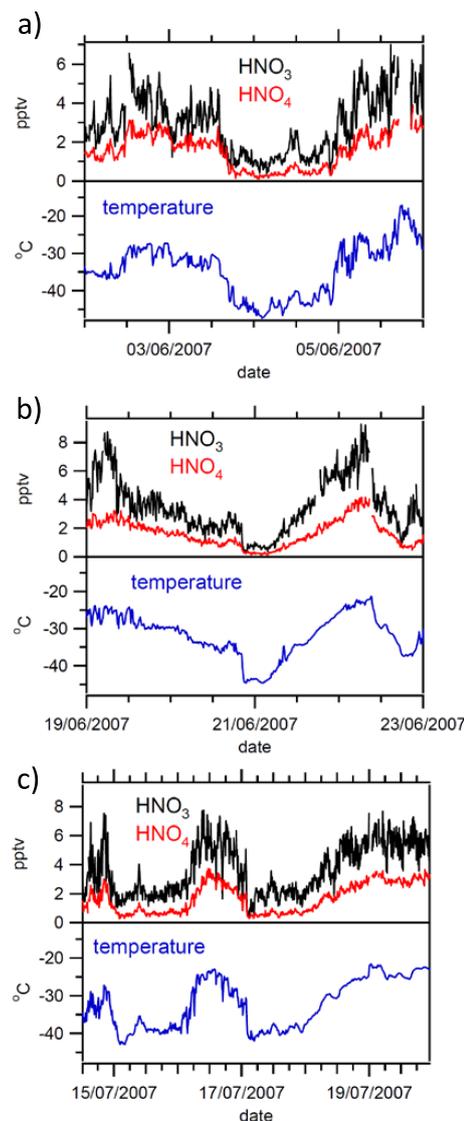
Table 1. Mean and standard deviation of nitric acid and peroxyntic acid according to ambient air temperature range. The statistics were derived using only data above the 3 σ detection limit.

Ambient air temperature (°C)	Nitric acid			Peroxyntic acid		
	no. points	mean (pptv)	SD (pptv)	no. points	mean (pptv)	SD (pptv)
−45 to −47.4	24	0.96	0.12	0		
−40 to −45.9	579	1.36	0.42	472	0.58	0.14
−35 to −39.9	1575	2.13	0.74	1615	0.99	0.38
−30 to −34.9	3638	3.09	0.92	3732	1.6	0.44
−25 to −29.9	4821	4.32	1.18	4942	2.29	0.62
−21.2 to −24.9	1954	5.67	1.19	2013	3.23	0.62

proach, we minimise any influence that air flow through the snow (e.g. via ventilation/wind pumping) may have on air/snow exchange processes. We derive a mixing diffusivity to determine the timescale for vertical mixing (via turbulent diffusion) between the snow surface and the CIMS inlet height, in order to confirm that the CIMS HNO₃ and HO₂NO₂ observations can be used to analyse processes occurring at the ground-level air/snow interface. The mixing diffusivity is roughly equal to $kz u_*$, where k is von Karman's constant (0.4), z is CIMS inlet height (5 m) and u_* is the friction velocity, derived from the sonic anemometer data (Stull, 1988). The e -folding time scale, t_{sc} , is given by $z^2/\text{diffusivity}$, that is $t_{sc} = z/(k u_*)$. During each case study t_{sc} will vary, but cannot be negative: this range is presented below derived from log means and standard deviations. Figure 6a shows observations made on 30 May 2007, with a clear gradual increase in both HNO₃ and HO₂NO₂ as ambient temperatures rose from ~ -44 to ~ -30 °C. On this day, data from the boundary layer mast (not shown) show that between the surface and 8 m height, there was little or no temperature gradient; to first order, therefore, 8 m temperatures can be used as a surrogate for those at the ground. Wind speeds were between 0 and 2 m s^{−1} from the surface to 4 m, and remained at around 2 m s^{−1} at 8 m height. For 95 % of the time, t_{sc} was between 100 and 300 s.

Figure 6b shows data for the period from 9 a.m. to midnight on 21 June, discussed briefly in Sect. 3.2 above. Again, a gradual increase in mixing ratios of HNO₃ and HO₂NO₂ is evident (upper panel), as ambient temperatures rose gradually from ~ -38 to ~ -26 °C. Data from the boundary layer met mast show that, during this period, there was no vertical gradient in temperature between the surface and 8 m height; wind speeds from the surface to 8 m were below 2 m s^{−1}. Data from the sonic anemometers show that vertical mixing was again very weak, with mixing time scales between 90 and 600 s.

Figure 6c shows observations from midnight to 9 a.m. on 18 July, another quiescent period, with wind speeds in the lowest 8 m below 1 m s^{−1}, and with no temperature gradient below 32 m. Gradual increases in both HNO₃ and HO₂NO₂

**Figure 5.** Three examples of short-term variability in HNO₃, HO₂NO₂ and ambient air temperature (10 min data), from (a) early June; (b) mid June; and (c) mid July. All three periods are during the winter 24 h per day darkness.

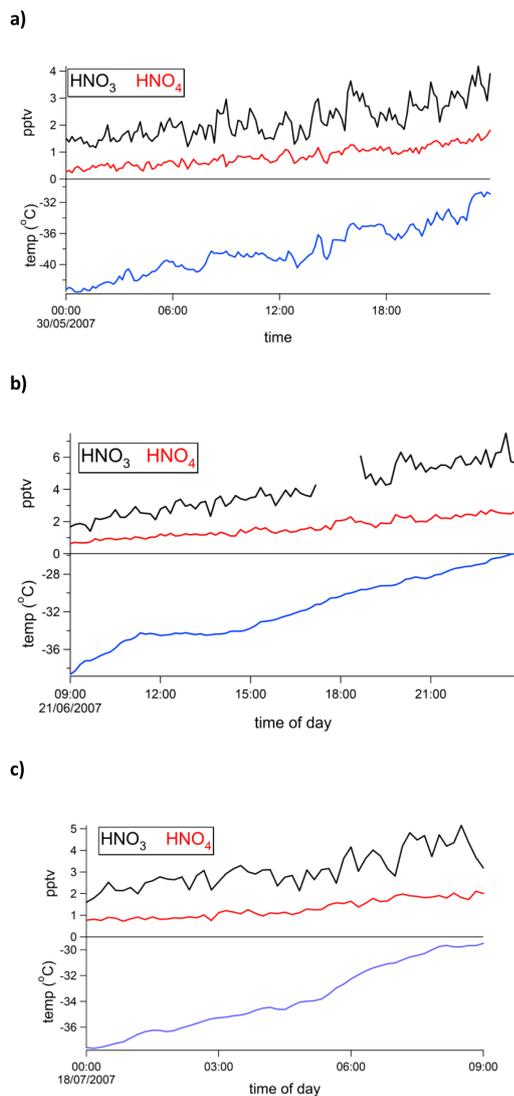


Figure 6. Detail of changes in HNO₃, HO₂NO₂ and temperature on (a) 30 May, (b) 21 June, and (c) 18 July. These three periods in the measurement series were characterised by low and invariant wind speeds and 24 h per day darkness.

proceed as ambient temperatures rise from -38 to -30 °C. t_{sc} varied between 30 and 100 s during the event.

Correlation coefficients between mixing ratios of HNO₃ (and HO₂NO₂) and temperature, are extremely high for the time periods presented in Fig. 6: R^2 for the correlation between HNO₃ and temperature is 0.72 (30 May), 0.90 (21 June) and 0.72 (18 July); for the correlation between HO₂NO₂ and temperature, R^2 was 0.88 (30 May), 0.94 (21 June), and 0.92 (18 July). The values of R^2 show that between 72 and 90 % of the variability in HNO₃ can be explained by variability in temperature; and between 88 and 94 % of the variability in HO₂NO₂ can be explained by the variability in temperature.

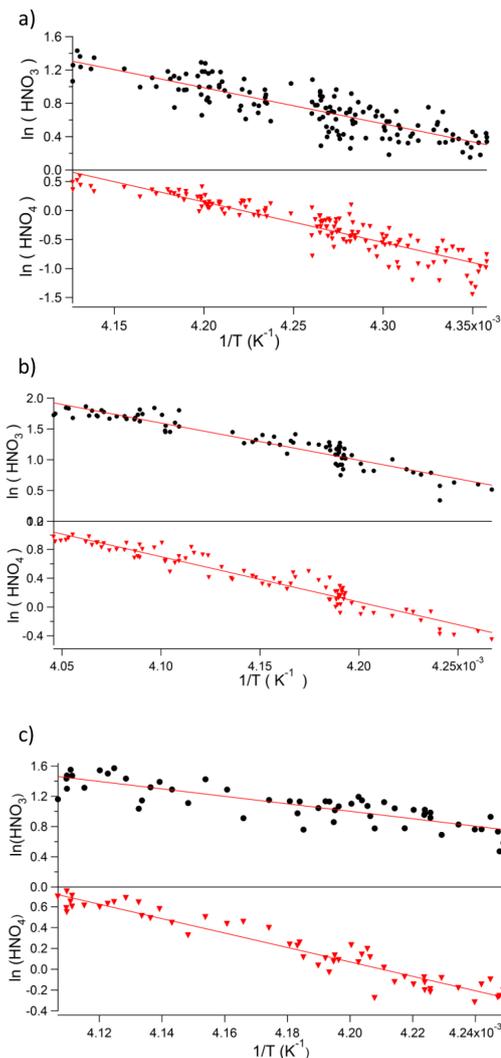


Figure 7. Plots of $\ln(\text{HNO}_3)$ and $\ln(\text{HO}_2\text{NO}_2)$ vs. $1/T$ for the time periods shown in Fig. 6, i.e. (a) 30 May, (b) 21 June, and (c) 18 July.

3.4 Deriving enthalpy of adsorption from the Halley field data

The enthalpies of adsorption between HNO₃/ice and HO₂NO₂/ice have been derived in laboratory experiments carried out under environmentally relevant conditions. Ulrich et al. (2012) studied uptake of HO₂NO₂ at low concentrations and temperatures between 230 and 253 K while Bartels-Rausch et al. (2002) and Ullerstam et al. (2005) studied the adsorption enthalpy of HNO₃. Field studies carried out during the 24 h per day darkness of Antarctic winter provide optimum conditions for validating such laboratory-derived physical air/snow exchange parameters under “real-world” conditions.

At equilibrium, the partitioning of HO₂NO₂ or HNO₃ molecules between the gas phase (C_g) and the snow/ice sur-

Table 2. The enthalpies of adsorption to ice for HO₂NO₂ and HNO₃ as derived from three periods of the Halley measurements. The number of observations used for each derivation is also given.

	HO ₂ NO ₂		HNO ₃	
	ΔH_{ads} (kJ mol ⁻¹)	no. observations	ΔH_{ads} (kJ mol ⁻¹)	no. observations
30 May	-58 ± 2	144	-36 ± 2	144
21 June	-52 ± 2	90	-50 ± 2	82
18 July	-58 ± 2	55	-41 ± 4	55

face (C_s) can be expressed as

$$K_{\text{part}} = \frac{C_s}{C_g}$$

As the partition constant will obey the van't Hoff equation, a new equation can be written as follows:

$$\frac{d \ln K_{\text{part}}}{d \frac{1}{T}} = -\frac{\Delta H_{\text{ads}}}{R},$$

where T is the temperature (K), ΔH is the enthalpy of adsorption (J mol⁻¹), R is the gas constant (8.314 J K⁻¹ mol⁻¹).

Given our definition of K_{part} above, the equation can then be re-formulated as

$$\frac{d \ln \frac{C_s}{C_g}}{d \frac{1}{T}} = -\frac{\Delta H_{\text{ads}}}{R}.$$

At 240 K, roughly the temperatures of our observations, Ulrich et al. (2012), in their Fig. 4, show $C_s / C_g \approx 20$ cm for HO₂NO₂ and 8000 cm for HNO₃. In the firn, the ratio of the surface area of snow to the volume of air is approximately 50 to 500 cm⁻¹ (based on a density of snow of 0.3 g cm⁻³ and a specific surface area of 100 to 1000 cm² g⁻¹; Domine et al., 2008). We can therefore calculate that the ratio of the number of molecules of HNO₃ adsorbed to the snow surface to that in the gas phase, is approximately 4×10^5 to 4×10^6 ; for HO₂NO₂, this ratio is 1000 to 10 000. As a result, for both HNO₃ and HO₂NO₂, exchange between the air and snow will thus have little effect on C_s , which can therefore be considered as a constant relative to C_g .

If we also assume that, over several hours, and under low and constant wind conditions, the concentration at our inlet tracks the concentration in the firn, then

$$\frac{d \ln C_g}{d \frac{1}{T}} = \frac{\Delta H_{\text{ads}}}{R}.$$

It is then possible to derive ΔH_{ads} from the slope of $\ln C_g$ vs. $1/T$ multiplied by R .

For the Halley data, Fig. 7 shows plots of both $\ln(\text{HNO}_3)$ vs. $1/T$ and $\ln(\text{HO}_2\text{NO}_2)$ vs. $1/T$ for the time periods discussed in Sect. 3.3 above. As a reminder, these periods are

characterised by 24 h per day darkness, low wind speeds, and limited vertical mixing from turbulent diffusion, so are as close to laboratory conditions as could be found in our data set. They were also chosen as they spanned a reasonably large temperature range, which would improve the constraint on the linear fit.

The values of ΔH_{ads} derived from these fits are given in Table 2. The average ΔH_{ads} for HNO₃ is -42 ± 2 kJ mol⁻¹ which can be compared with laboratory-derived values of Bartels-Rausch et al. (2002) and Ullerstam et al. (2005). Bartels-Rausch et al. (2002) derived ΔH_{ads} of -44 kJ mol⁻¹ (with random error 2.3 kJ mol⁻¹, systematic error 13 kJ mol⁻¹); Ullerstam et al. (2005), working at lower concentrations of HNO₃, relevant to the natural atmosphere, derived ΔH_{ads} of -30.6 ± 6.0 kJ mol⁻¹. For HO₂NO₂, the average ΔH_{ads} derived from our field data is -56 ± 1 kJ mol⁻¹ which can be compared with the laboratory-derived value (Ulrich et al., 2012) of -59 kJ mol⁻¹. For both HO₂NO₂ and HNO₃, the agreement between laboratory and field-derived enthalpies of adsorption is remarkably good.

4 Summary and conclusions

We present the first high time resolution observations of HNO₃ and HO₂NO₂ in coastal Antarctica, and the first from Antarctica during the dark winter period. Mixing ratios of HNO₃ ranged from instrumental detection limits to ~ 8 parts per trillion by volume (pptv) and of HO₂NO₂ varied from detection limits to ~ 5 pptv. These values are on average lower than those observed at the South Pole in summer, where mixing ratios of HNO₃ and HO₂NO₂ were generally in the 10s of pptv, and sometimes over 100 pptv.

The Antarctic, during winter, is an ideal natural laboratory for studying physical air/snow exchange processes. The environmental system is considerably simplified compared with other times of the year because of the lack of photochemical activity which must otherwise be taken into account when interpreting data.

In our study, we considered whether adsorption/desorption of HNO₃ and HO₂NO₂ to snow/ice surfaces could be invoked to explain our observations; we did not consider formation of solid solutions from solid ice, or take-up to liquid

NaCl aerosols. In a follow-up study that further analysed our field data, Bartels-Rausch (2014) considered both the case of Langmuir adsorption to the ice surface, and solubility in ice forming a solid solution. He found that equilibrium air/snow partitioning was able to describe our field data well, both in terms of absolute mixing ratios and trend with temperature. He also found that the reservoir of adsorbed HNO₃ and HO₂NO₂ in the upper snow pack was sufficient to fuel the observed emissions. In contrast, while calculations based on reversible solid-solution/air partitioning were able to describe mixing ratios of HNO₃, they were not able to reproduce the observed trend with temperature. Further, the reservoir of HNO₃ in the outer part of the snow crystals was too small to explain observed increases in mixing ratio.

The measurements of HNO₃ and HO₂NO₂ from Halley are consistent with laboratory experiments showing a temperature-dependence in the partitioning of both HNO₃ and HO₂NO₂ to ice. They further support the conclusion that HO₂NO₂/ice interactions are stronger than those between HNO₃ and ice, as shown by the higher enthalpy of adsorption of HO₂NO₂ compared with HNO₃ (Ulrich et al., 2012). On short timescales, therefore, HNO₃ and HO₂NO₂ that is adsorbed to snow/ice can be re-released as temperatures rise. The snowpack can thus act as a source of HNO₃ and HO₂NO₂ to the overlying atmosphere at all times of the year given sufficient reservoir in the snowpack and changing temperatures. Similarly, HNO₃ and HO₂NO₂ adsorbed to cirrus clouds would be desorbed should temperatures rise. Such a reversible, temperature-dependent partitioning also provides a mechanism for re-distributing HNO₃ and HO₂NO₂ on a local or regional scale across Antarctica. Snow can be transported considerable distances by storm systems, and adsorbed HO₂NO₂ and HNO₃ can be desorbed as a function of changing temperature experienced along the transport pathway. Indeed, transport of snow from inland Antarctica is likely to contribute to the HO₂NO₂ and HNO₃ reservoir in the coastal snowpack. The other likely source for HO₂NO₂ to the winter snowpack is the general shift in equilibria, as temperatures fall from summer to winter, from gas-phase HO₂ and NO₂ towards gaseous HO₂NO₂, and then, by temperature-dependent partitioning, towards snowpack-adsorbed HO₂NO₂.

While clearly a controlling mechanism during polar night, the importance of air/snow partitioning relative to photochemistry will vary according to time of year and location. However, adsorption to/desorption from the snow pack should be taken into account when addressing budgets of boundary layer HO₂NO₂ and HNO₃ at any snow-covered site, as all are likely to experience varying ambient temperature which would drive such air/snow exchange.

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