



# Global and regional impacts of HONO on the chemical composition of clouds and aerosols

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**Abstract.** Recently, realistic simulation of nitrous acid (HONO) based on the HONO/NO<sub>x</sub> ratio of 0.02 was found to have a significant impact on the global budgets of HO<sub>x</sub> (OH + HO<sub>2</sub>) and gas phase oxidation products in polluted regions, especially in winter when other photolytic sources are of minor importance. It has been reported that chemistry-transport models underestimate sulphate concentrations, mostly during winter. Here we show that simulating realistic HONO levels can significantly enhance aerosol sulphate (S(VI)) due to the increased formation of H<sub>2</sub>SO<sub>4</sub>. Even though in-cloud aqueous phase oxidation of dissolved SO<sub>2</sub> (S(IV)) is the main source of S(VI), it appears that HONO related enhancement of H<sub>2</sub>O<sub>2</sub> does not significantly affect sulphate because of the predominantly S(IV) limited conditions, except over eastern Asia. Nitrate is also increased via enhanced gaseous HNO<sub>3</sub> formation and N<sub>2</sub>O<sub>5</sub> hydrolysis on aerosol particles. Ammonium nitrate is enhanced in ammonia-rich regions but not under ammonia-limited conditions. Furthermore, particle number concentrations are also higher, accompanied by the transfer from hydrophobic to hydrophilic aerosol modes. This implies a significant impact on the particle lifetime and cloud nucleating properties. The HONO induced enhancements of all species studied are relatively strong in winter though negligible in summer. Simulating realistic HONO levels is found to improve the model-measurement agreement of sulphate aerosols, most apparent over the US. Our results underscore the importance of HONO for the atmospheric oxidizing capacity and corroborate the central role of cloud chemical processing in S(IV) formation.

## 1 Introduction

Despite improvements in the control strategies of air quality in urban areas during the last decades, reducing air pollution is still a major challenge due to the very complex chemical mechanisms and the large number of species emitted into the atmosphere. Fine aerosol particles have direct and indirect impacts on climate (e.g., Charlson et al., 1991, 1992; Haywood and Boucher, 2000; IPCC, 2007) and human health (e.g., WHO, 2002; Lelieveld et al., 2013). The aerosol radiative forcing is highly uncertain and depends on the aerosol burden and distribution (IPCC, 2007). Knowledge of aerosol chemical composition and the controlling processes are thus of paramount importance. The inorganic compounds SO<sub>4</sub><sup>2-</sup>, NO<sub>3</sub><sup>-</sup> and NH<sub>4</sub><sup>+</sup> are among the major constituents of the fine particles in polluted air, e.g., in megacities such as Beijing (Sun et al., 2013). Sulphate aerosols are primarily produced by the condensation of H<sub>2</sub>SO<sub>4</sub> and the aqueous phase oxidation of dissolved gas phase SO<sub>2</sub>, S(IV), by O<sub>3</sub> and H<sub>2</sub>O<sub>2</sub> (e.g., Lelieveld et al., 1997; Seinfeld and Pandis, 2006). Recently, the transition-metal-ion catalysed oxidation of S(IV) by O<sub>2</sub> has also been highlighted as an important source especially in rural areas (Harris et al., 2013, and references therein). Cloud processing has long been known as the most important source of aerosol sulphate on regional (e.g., Seidl and Dlugi, 1991; McHenry and Dennis, 1994; Lelieveld et al., 1997; Venkataraman et al., 2001; Menegoz et al., 2009) and global scale (e.g., Hegg, 1985; Feichter et al., 1996; Liao et al., 2003), accounting for up to two thirds of ambient sulphate concentrations, with H<sub>2</sub>O<sub>2</sub> being the most effective oxidant followed by O<sub>3</sub> (e.g., Feichter et al., 1996).

The aqueous phase oxidation of S(IV) can be either oxidant (i.e., O<sub>3</sub> or H<sub>2</sub>O<sub>2</sub>) or SO<sub>2</sub> limited and occurs primarily in clouds because of their relatively high liquid water content, LWC, compared to aerosol (e.g., Schwartz et al., 1987; Van den Berg et al., 2000; Seinfeld and Pandis, 2006). Under oxidant limited conditions, enhancement of the oxidant concentrations may significantly accelerate sulphate formation, in contrast to S(IV) limited conditions (e.g., Seinfeld and Pandis, 2006). Secondary aerosol formation by gas/liquid partitioning is also well recognised as an important source of aerosols (e.g., Kulmala et al., 2004). Sulphuric acid is the major contributor to atmospheric nucleation (e.g., Seidl and Dlugi, 1991; Kulmala et al., 2006; Kirkby et al., 2011; Neitola et al., 2013). Unlike sulphates, nitrates are nearly exclusively produced by the uptake of nitric acid (HNO<sub>3</sub>) into cloud droplets and aerosols and its reaction with available ammonium and other cations. In ammonia-rich regions, enhancement of aerosol sulphate concentrations may also enhance nitrate aerosol formation via the heterogeneous N<sub>2</sub>O<sub>5</sub> hydrolysis on sulphate aerosol particles forming HNO<sub>3</sub> (e.g., Lamsal et al., 2010; L. Zhang et al., 2012). In ammonia-limited regions, enhancements of sulphate can substitute nitrate and chloride ions, driving them out of the condensed phase (e.g., Seinfeld and Pandis, 2006).

During the last decades, HONO has been shown to play a major role in HO<sub>x</sub> (OH + HO<sub>2</sub>) and O<sub>3</sub> production as well as in the formation of other secondary oxidation products (e.g., Perner and Platt, 1979; Harris et al., 1982; Elshorbany et al., 2009a, b, 2010a, b, 2012a, b). Several previous local-scale studies have reported an important impact of HONO photolysis on aerosol chemical composition. Goncalves et al. (2012) found that simulated PM<sub>2.5</sub> is increased up to 14 % by accounting for HONO sources in Madrid, Spain. Li et al. (2010) found that HONO sources play an important role in the formation of secondary aerosols in Mexico City, substantially enhancing their concentrations by a factor of 2 on average in the morning. They also found that the simulated particle-phase nitrate and ammonium are substantially enhanced in the morning, though the effect on sulphate aerosol was much smaller, being in good agreement with the measurements. Further, R. Zhang et al. (2012) found that HONO can enhance the daily mean PM<sub>2.5</sub> in the Pearl River Delta, China by up to 17 μg m<sup>-3</sup> (12 %).

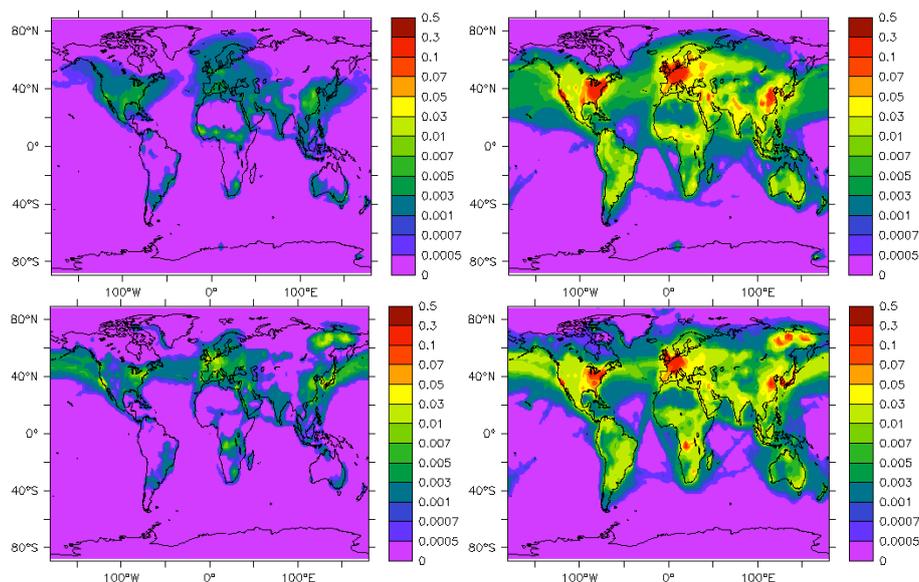
Underestimation of sulphate aerosols has been reported by several studies, especially in polluted regions, which was attributed to missing oxidation pathways, being largest in winter and relatively small in summer (e.g., Dennis et al., 1993; Roelofs et al., 1998, 2001). HONO formation mechanisms are still poorly understood and therefore it is difficult for global models to simulate realistic HONO concentrations. Recently, Elshorbany et al. (2012b) have shown that HONO levels can be realistically parameterized, being about an order of magnitude higher compared to the reference simulations that consider the reaction of OH + NO as the sole HONO source. They also showed that HONO photolysis sig-

nificantly enhances the gas phase mixing ratios of H<sub>2</sub>O<sub>2</sub>, HNO<sub>3</sub>, O<sub>3</sub> and H<sub>2</sub>SO<sub>4</sub>, which play an important role in the formation of aerosols. Therefore, simulating realistic HONO levels may possibly improve the agreement between models and measurements. In this study, we investigate the impact of HONO on the aerosol physical and chemical properties and compare the simulation results with measurement data from monitoring networks.

## 2 Model description

The applied modelling system based on the ECHAM5 general circulation model (Roeckner et al., 2006) and the Modular Earth Submodel System (MESSy2, Jöckel et al., 2005, 2010) to simulate the meteorology and atmospheric chemistry. The ECHAM5/MESSy Atmospheric Chemistry (EMAC) system is a coupled lower-middle atmospheric chemistry general circulation model (AC-GCM), which has been extensively evaluated against observations (e.g., Jöckel et al., 2006; Lelieveld et al., 2007; Tost et al., 2007; Pozzer et al., 2007; 2010, 2012; Brühl et al., 2012; de Meij et al., 2012). The model structure and setup have been described by Jöckel et al. (2006) and (2010) and only a brief description is given here, focusing on aerosol related submodels.

The atmospheric aerosol is parameterized in this study using the submodel GMXe (Pringle et al., 2010; Tost and Pringle, 2012). Gas/aerosol partitioning is calculated in GMXe in two stages; *kinetically* (the same as that in M7 Vignati et al., 2004, but extended to NH<sub>3</sub>, HCl and HNO<sub>3</sub>), in which the amount of gas phase species is kinetically allowed to condense (assuming diffusion limited condensation) on aerosol particles and *thermodynamically* (using ISORROPIA II, Fountoukis and Nenes, 2007), in which the condensed material is redistributed between the gas and aerosol phase. The size distribution (in radius) is described by four hydrophilic (nucleation (ns, < 5 nm), Aitken (ks, 5–50 nm), accumulation (as, 50–500 nm), coarse (cs, > 500 nm)) and three hydrophobic (Aitken (ki, 5–50 nm), accumulation (ai, 50–500 nm), coarse (ci, > 500 nm)) interacting lognormal aerosol modes, with prognostic aerosol mass and number and a diagnosed mean radius. The aerosol compounds treated explicitly are H<sub>2</sub>O, Na<sup>+</sup>, Cl<sup>-</sup>, NH<sub>4</sub><sup>+</sup>, SO<sub>4</sub><sup>2-</sup>, NO<sub>3</sub><sup>-</sup> and the bulk compounds of organic carbon (OC), black carbon (BC), dust and a bulk sea salt (the fraction that is not described by NaCl and sea salt sulphate). Dry deposition, sedimentation and wet deposition of gas and aerosol phase species are calculated using DRYDEP (dry deposition of trace gases and aerosol particles, Ganzeveld et al., 1998; Kerckweg et al., 2006), SEDI (sedimentation of aerosol particles, Kerckweg et al., 2006), and SCAV (scavenging and liquid phase chemistry in clouds and precipitation, Tost et al., 2006) submodels, respectively. SCAV calculates the uptake of trace gases and aerosol particles (radius dependent) into cloud and precipitation droplets. The wet deposition flux out of the lowest

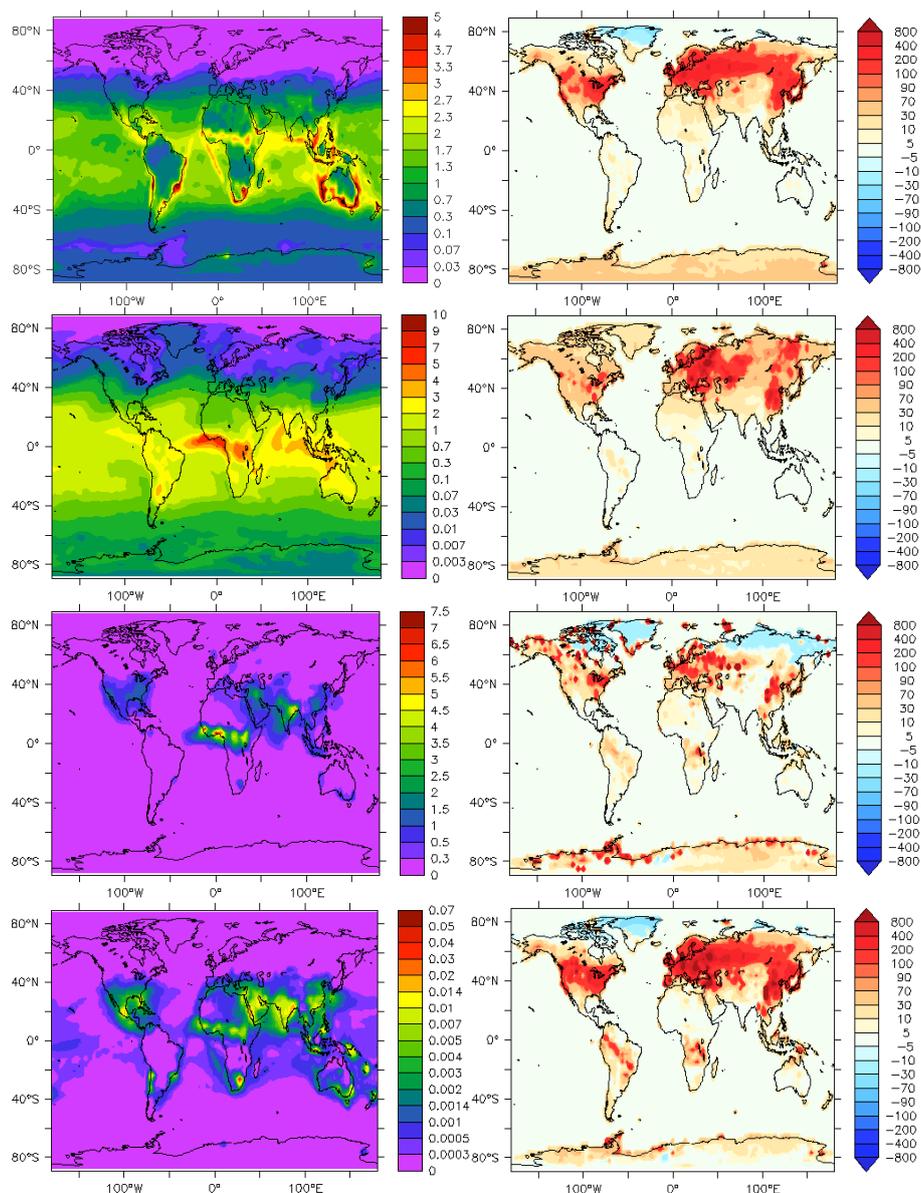


**Fig. 1.** Simulated monthly average HONO mixing ratios (ppbv) from the reference run (base\_S1, left) and the sensitivity run (S1, i.e. using a HONO/NO<sub>x</sub> ratio of 0.02, right) near the surface in January (upper panels) and July (lower panels).

model layer represents the chemical composition of rainwater. The gas and aqueous phase chemistry, including the transfer of gaseous compounds, dissociation of acidic and alkaline species in the droplets and aqueous phase redox reactions are calculated by a coupled system of ordinary differential equations using the kinetic pre-processor (KPP) software (Sandu and Sander, 2006). The gas and the liquid phase chemical processes are fully coupled and do not require prescribed mixing ratios or pH values. Since usually only a small fraction of a grid box is affected by clouds and precipitation, only the cloud covered part or the part in which the precipitation occurs contributes to the scavenging, while the rest of the grid box remains unaffected. A scavenged aerosol particle can be either removed from the atmosphere by wet deposition or released through droplet evaporation (aerosols are then transferred into the largest available mode). Ions that are converted to molecules via chemical reactions in the aqueous phase can be released to the gas phase. At the end of each time step, it is assumed that the cloud completely evaporates and that all volatile species will be released to the gas phase, while the ions are transferred to the aerosol phase, thus affecting the aerosol properties: e.g., the SO<sub>2</sub> oxidation to SO<sub>4</sub><sup>2-</sup> can significantly increase the aerosol sulphate amount. Atmospheric gas phase chemical reactions are incorporated in the model through the module MECCA (Sander et al., 2005, 2011). To study the effects of changes in the chemistry while avoiding possible feedbacks of radiatively active gases and aerosols through the meteorology, the radiation scheme has been decoupled, and the model is used in the atmospheric chemistry-transport mode. The radiation code in EMAC uses an ozone climatology (Fortuin and Kelder, 1998), fixed ver-

tical profiles for CH<sub>4</sub>, N<sub>2</sub>O and CFCs and constant mixing ratios of CO<sub>2</sub>.

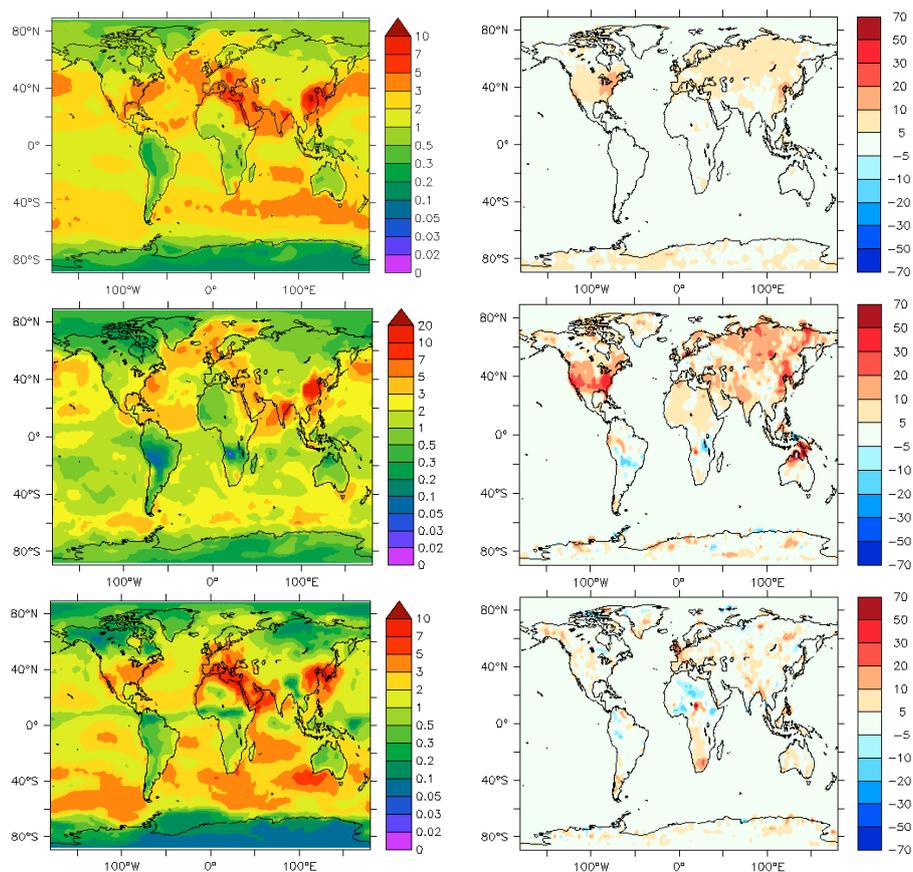
HONO is parameterized in this study following Elshorbany et al. (2012b), in which HONO, NO<sub>x</sub> and auxiliary atmospheric chemistry parameters have been investigated using data from 15 field measurement campaigns around the globe. The high correlation between HONO and NO<sub>x</sub> in all data sets revealed a robust and consistent linear regression slope of  $0.02 \pm 0.002$ . Comparison with field measurements showed that, given the ambient NO<sub>x</sub> mixing ratio, the HONO/NO<sub>x</sub> ratio is a suitable proxy predictor for HONO mixing ratios under different atmospheric conditions. Using a global chemistry-climate model and employing this HONO/NO<sub>x</sub> ratio, realistic HONO levels are simulated, being about one order of magnitude higher than the reference calculations that only consider the reaction  $\text{OH} + \text{NO} \rightarrow \text{HONO}$ . The resulting enhancement of HONO was found to significantly impact HO<sub>x</sub> levels and photo-oxidation products (e.g. O<sub>3</sub>, PAN), mainly in polluted regions. Furthermore, the relative enhancements in OH and secondary products were higher in winter than in summer, thus enhancing the oxidation capacity in polluted regions, especially in winter when other photolytic OH sources are of minor importance. Similar to Elshorbany et al. (2012b), HONO levels are parameterized in this study (as mentioned in the sensitivity model runs below) such that the HONO/NO<sub>x</sub> ratio is fixed at 0.02, which was achieved by an iterative correction of HONO and NO<sub>x</sub> (i.e., without perturbing the reactive nitrogen budget), which was applied every minute, thus limiting deviations in the HONO/NO<sub>x</sub> ratio to within  $\pm 0.005$ .



**Fig. 2.** Average simulated OH (from top to bottom) ( $10^6$  molecules  $\text{cm}^{-3}$ ),  $\text{H}_2\text{O}_2$ ,  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$  (ppbv) near the surface from the reference run (base\_S1, left panels) and their relative enhancements (%) due to the implementation of realistic HONO levels in the model (S1, right panels) in January.

The reference run (base\_S1) has been performed for the years 2000–2001 in T42L31 resolution (i.e., with a triangular truncation at wave number 42 for the spectral core of ECHAM5, and with 31 levels on a hybrid-pressure grid in the vertical, reaching up to 10 hPa). The T42 resolution corresponds to a quadratic Gaussian grid of approximately  $2.8^\circ \times 2.8^\circ$  in latitude and longitude. To simulate realistic synoptic conditions, we applied a weak “nudging” towards actual meteorology by the assimilation of analysis data from the European Centre for Medium-range Weather Forecasting (ECMWF) through the Newtonian relaxation of four prognostic model variables: temperature, divergence, vorticity

and the logarithm of surface pressure (van Aalst et al., 2004; Lelieveld et al., 2007). The first 12 months of the simulation are regarded as a spin up period and these results are not considered in our analysis. In the sensitivity run S1, HONO levels were parameterized such that the HONO/ $\text{NO}_x$  ratio is fixed at 0.02, simulating realistic HONO levels (Elshorbany et al., 2012b). In order to investigate the most important process affected by the HONO enhancement in the model, four additional sensitivity runs were performed (see Table 1) base\_S2, S2, base\_S3 and S3. In the base\_S2 and S2 sensitivity runs, only gas/liquid partitioning based on Henry’s law coefficients was allowed without aqueous phase chemistry



**Fig. 3.** Simulated aerosol sulphate concentrations ( $\mu\text{g m}^{-3}$ ) near the surface for the mean annual (upper panels), boreal winter (January, middle panels) and summer (July, lower panels) conditions from the reference run (base\_S1) and relative changes (in %, right panels) due to HONO enhancement in the model.

(similar to the “EASY” sensitivity runs in Tost et al., 2007), thus determining the impact of aqueous phase chemistry on aerosol formation in the absence (base\_S2) and the presence (S2) of realistic HONO levels by comparing it to the base\_S1 and S1 runs, respectively. In the base\_S3 and S3 sensitivity runs, the gas phase oxidation of  $\text{SO}_2$  with OH results in a dummy species (i.e.,  $\text{OH} + \text{SO}_2 \rightarrow \text{DUMMY}$ ), thus masking the production of gas phase  $\text{H}_2\text{SO}_4$ . From these sensitivity runs, the impact of gas phase  $\text{H}_2\text{SO}_4$  on the aerosol formation in the absence (base\_S3) and the presence (S3) of realistic HONO levels can be determined by comparing them to the base\_S1 and S1 runs, respectively.

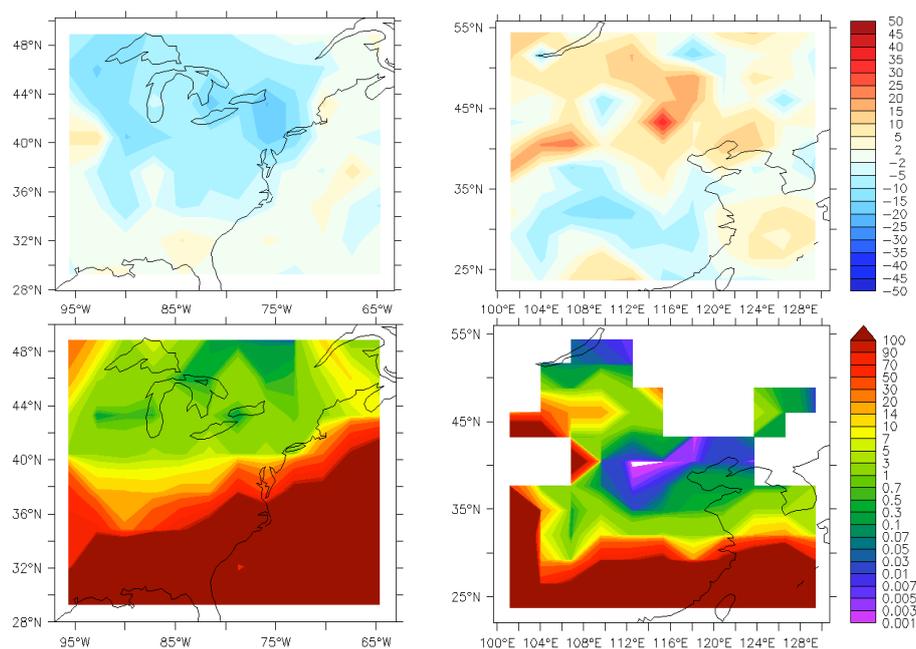
The current study builds on the model-observation results of Elshorbany et al. (2012b) for the HONO parameterization and on the aerosol validation results of Pringle et al. (2010), Tost and Pringle (2012), de Meij et al. (2012) and Pozzer et al. (2012). Furthermore, modelling results of this study are also compared to data from monitoring networks (see Sect. 4.2 and Table 5).

### 3 Measurement networks

The simulated aerosol concentrations are compared with data from the Clean Air Status and Trends NETWORK (CASTNET; Edgerton et al., 1990) for North America, the European Monitoring and Evaluation Program (EMEP; Hjellbrekke and Fjæraa, 2011) for Europe and the Acid Deposition Monitoring Network in East Asia (EANET; Totsuka et al., 2005) for Asia. The model result – measurements comparisons are performed for the simulated period in 2001, on a monthly mean basis. Only stations that cover the entire time period within each network are included, resulting in 54, 81 and 10 stations for CASTNET, EMEP and EANET, respectively. For comparison, the simulated data are sampled at the observation locations based on the respective geographical coordinates of the measuring stations.

### 4 Results and discussion

As shown in Fig. 1, the monthly mean simulated near surface HONO levels during both summer and winter are about one



**Fig. 4.** Simulated monthly average relative changes (%) of aerosol sulphate (upper panels), and the ratio of in-cloud aqueous phase concentrations of  $\text{H}_2\text{O}_2/\text{HSO}_3^-$  (lower panels) in January without the gas phase oxidation of  $\text{SO}_2$  (base\_S3), over the eastern US (left) and China (right).

order of magnitude higher in the S1 run than in the base\_S1 run, which considers the gas phase reaction of NO with OH as the sole source of HONO, in line with our previous analysis (Elshorbany et al., 2012b). Elshorbany et al. (2012b) have shown that simulating realistic HONO levels significantly enhances the global concentrations of  $\text{HO}_x$  and oxidation products in high  $\text{NO}_x$  regions, especially during the winter season when other photolytic sources are of minor importance. In the next sections, we focus on the HONO impacts during the boreal winter season, while the summer and/or annual results are presented in the Supplement. In Fig. 2, the impact of the enhanced HONO levels on the simulated monthly average OH and the gas phase oxidation products,  $\text{HNO}_3$  ( $\text{OH} + \text{NO}_2$ ),  $\text{H}_2\text{SO}_4$  ( $\text{OH} + \text{SO}_2$ ) and  $\text{H}_2\text{O}_2$  ( $\text{HO}_2 + \text{HO}_2$ ) during winter are shown. The enhancement of HONO leads to increased oxidation rates by OH leading to enhancements of  $\text{O}_3$ ,  $\text{N}_2\text{O}_5$  (not shown) and  $\text{H}_2\text{O}_2$  and also  $\text{HNO}_3$  and  $\text{H}_2\text{SO}_4$  levels, especially during winter (Fig. 2) while in summer much smaller changes are simulated. The enhancements of  $\text{HNO}_3$  and  $\text{H}_2\text{SO}_4$  may increase nitrate and sulphate aerosols, respectively. In addition, the enhancement of aerosol sulphate also contributes to the formation of nitrates via the heterogeneous hydrolysis of  $\text{N}_2\text{O}_5$  forming  $\text{HNO}_3$ . The uptake of  $\text{O}_3$  and  $\text{H}_2\text{O}_2$  in clouds may enhance the aqueous oxidation of dissolved  $\text{SO}_2$  in clouds, thus increasing the oxidation rates of  $\text{HSO}_3^-$  and  $\text{SO}_3^{2-}$  (S(IV)) to S(VI), which can be transferred to the aerosol phase upon the evaporation of the cloud droplets. However, the impact of HONO on the aqueous phase production of S(VI) depends

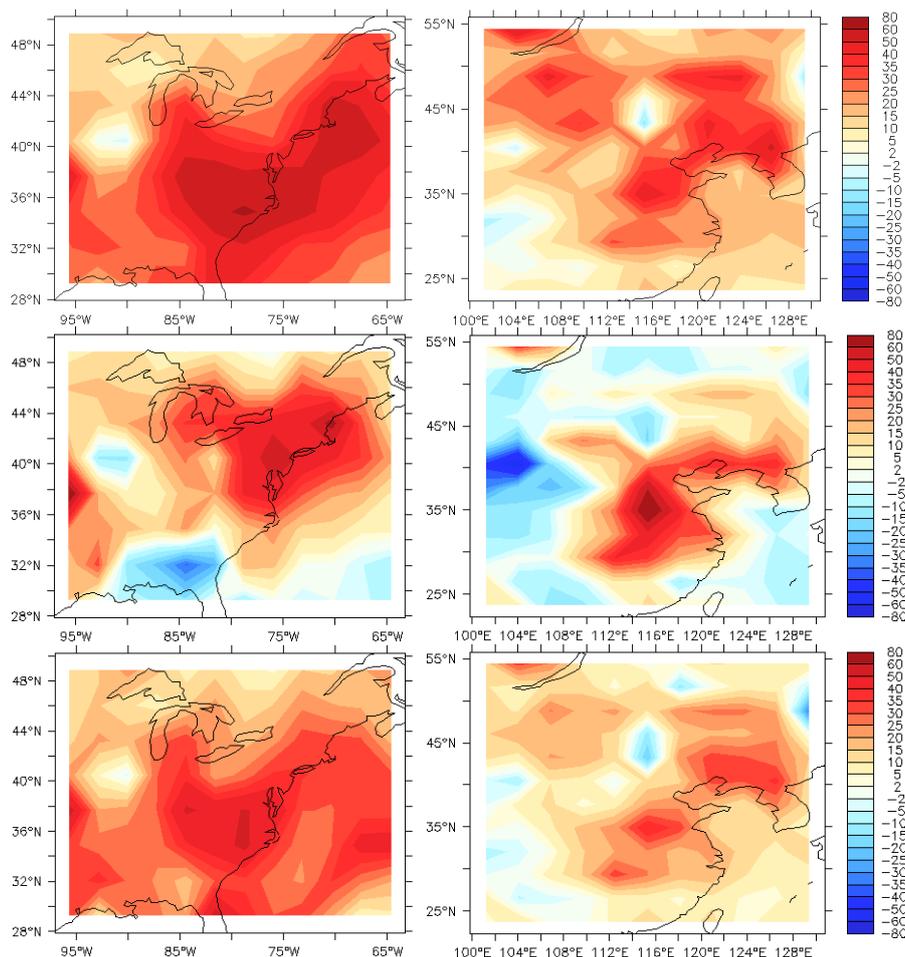
on the relative abundance of S(IV) and the oxidants and does not affect sulphate concentrations under S(IV) limited conditions. In the following sections, we investigate the impact of these enhanced gas phase species on the chemical composition of aerosols.

#### 4.1 Impacts on aerosol chemical composition

The major constituents of fine aerosol particles are typically organic compounds (lumped as OC), water soluble ions, e.g.,  $\text{SO}_4^{2-}$ ,  $\text{NO}_3^-$ , and  $\text{NH}_4^+$ , and mineral dust (Jimenez et al., 2009; Sun et al., 2013, and references therein). As mentioned above, daytime HONO photolysis accelerates oxidation processes, especially under high  $\text{NO}_x$  conditions, which leads to increased concentrations of aerosol precursors, e.g.,  $\text{HNO}_3$ ,  $\text{N}_2\text{O}_5$ ,  $\text{H}_2\text{SO}_4$  and other inorganic (e.g.,  $\text{O}_3$ ,  $\text{H}_2\text{O}_2$ ) and organic (OVOC, PAN) oxidation products, which can affect the cloud and aerosol composition.

##### 4.1.1 Sulphates

As shown in Fig. 3, sulphate concentrations are higher in summer than in winter, related to the faster photochemical oxidation during summer. On the other hand, the impact of HONO is strongest during winter related to the relatively minor importance of other photolytic OH sources and the higher NO levels (i.e., conditions under which OH recycling is most efficient, see Sect. 4, Elshorbany et al., 2010b, 2012b). Aerosol sulphates are enhanced only in regions that are rich in both  $\text{NO}_x$  and ammonia; i.e., though



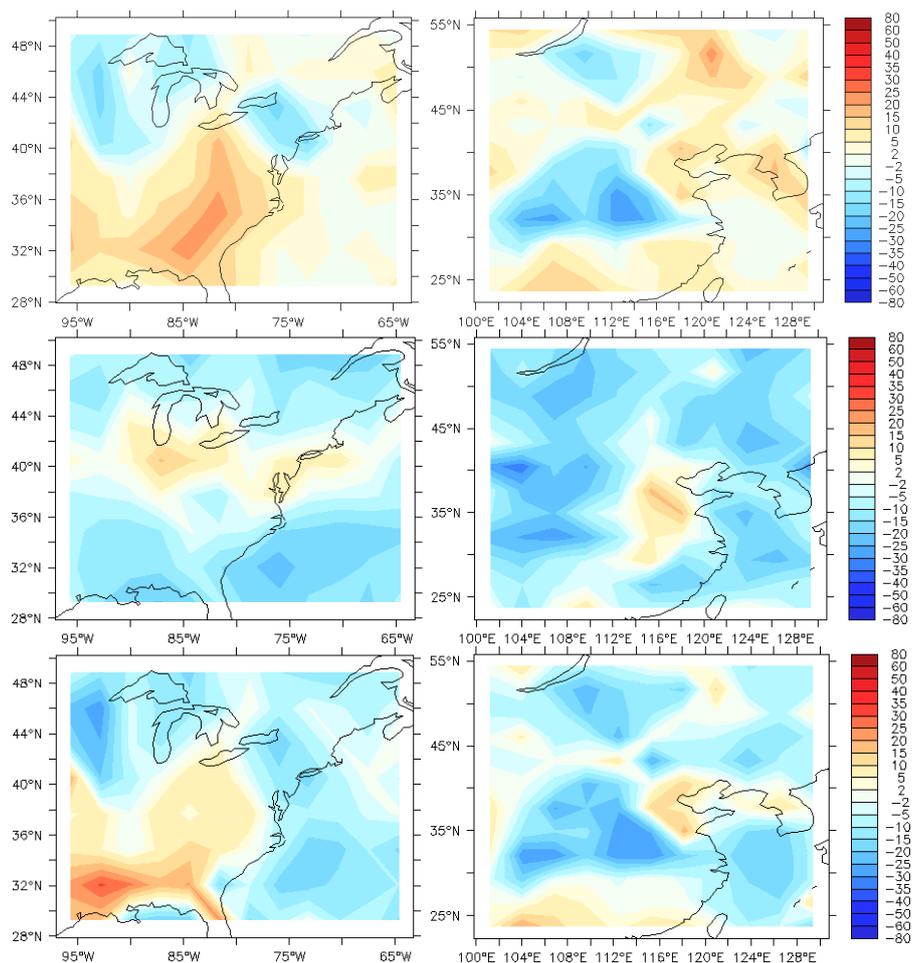
**Fig. 5.** Simulated monthly average relative changes (%) in the aerosol (accumulation soluble mode, as) sulphates (upper panels), nitrates (middle panels) and ammonium (lower panels) in the reference run (base\_S1) near the surface in January over the eastern US (left) and China (right) due to HONO enhancement in the model (S1).

India is an ammonia-rich region, sulphates are not significantly enhanced due to the relatively low- $\text{NO}_x$  conditions in this region (see Fig. 3). Considering the growing  $\text{NO}_x$  sources in southern Asia this may change in future. The mean annual aerosol concentrations of sulphates ( $\text{S(VI)}$ ) are significantly enhanced as a result of the more realistic HONO levels, reaching a mean annual relative enhancement of about 20 % over the highly polluted  $\text{NO}_x$ -rich regions, especially over the eastern US and China. Furthermore,  $\text{S(VI)}$  is significantly enhanced during winter, reaching up to +50 % during January (Fig. 3).

As shown in Table 2, gas phase  $\text{H}_2\text{SO}_4$  accounts for about 17–18 % of the simulated total sulphate aerosol concentration while aqueous phase chemistry accounts for about 25 and 54 % in January over the eastern US and China, respectively, in the base\_S1 run. The remainder of the sulphate concentration is due to emissions (i.e., direct anthropogenic emissions and sea salt sulphate, see Pringle et al., 2010; Tost and Pringle, 2012). As a result of simulating realistic HONO

levels (in the S1 run), the contribution of gas phase  $\text{H}_2\text{SO}_4$  increases to about 30 and 26 % while that of the aqueous chemistry decreases to 14 and 45 % over the eastern US and China, respectively. If only chemically formed sulphates are considered, gas phase  $\text{H}_2\text{SO}_4$  would account for about 41 and 25 % of the simulated sulphate aerosol concentrations, which increases to about 68 and 37 % in January over the eastern US and China, respectively, as a result of simulating realistic HONO levels (see Table 2). Thus, the relative contribution of gas phase  $\text{H}_2\text{SO}_4$  increases by about 69 and 50 % while that of the aqueous phase chemistry decreases by about 47 and 16 % over the eastern US and China, respectively. A striking result is that gas phase  $\text{H}_2\text{SO}_4$  becomes more important than aqueous phase chemistry in sulphate production over the eastern US (see Table 2).

In the absence of aqueous phase chemistry (Tables 2, S2), the enhancement of gas phase  $\text{H}_2\text{SO}_4$  leads to about 31 and 34 % enhancement over the eastern US and China, respectively. In the absence of gas phase  $\text{H}_2\text{SO}_4$  (Table 2,



**Fig. 6.** Simulated monthly average relative changes (%) in the aerosol (coarse soluble mode, cs) sulphates (upper panels), nitrates (middle panels) and ammonium (lower panels) in the reference run (base\_S1) near the surface in January over the eastern US (left) and China (right) due to HONO enhancement in the model (S1).

**Table 1.** List of the simulation setups of the different model scenarios.

| scenario | base_S   | S  |
|----------|--|--|
| 1        | default setup; [HONO]pss<br>(see Sect. 2)  | as base_S1 except HONO:<br>Iterative correction of HONO/NO <sub>x</sub> ratio = 0.02 |
| 2        | as base_S1 except SCAV = EASY: no aqueous phase chemistry,<br>i.e., only uptake of gas phase species using effective Henry's law constant          | as S1 except SCAV = EASY   |
| 3        | as base_S1 except SO <sub>2</sub> +OH = DUMMY,<br>i.e., no gas phase H <sub>2</sub> SO <sub>4</sub> from the oxidation of SO <sub>2</sub> with OH. | as S1 except SO <sub>2</sub> +OH = DUMMY   |

scenario S3), no significant sulphate enhancement is calculated. Thus, though both gas phase H<sub>2</sub>O<sub>2</sub> and O<sub>3</sub> are also enhanced as a result of simulating realistic HONO levels (Fig. 2), which subsequently increase their levels in the aqueous phase (see Supplement), they do not seem to significantly affect sulphate concentrations over the eastern US and China (see Table 2, scenario S3). This is related to the prevailing S(IV) limited conditions in most of these regions, thus any

increase in the H<sub>2</sub>O<sub>2</sub> concentration does not affect sulphate aerosol concentration (cf. Seinfeld and Pandis, 2006; Shen, 2011). For instance, over eastern China, in the absence of H<sub>2</sub>SO<sub>4</sub> from the gas phase oxidation of SO<sub>2</sub>, enhancement in the sulphate aerosol concentration (up to +50 %) seems to occur only when the ratio of H<sub>2</sub>O<sub>2</sub>/HSO<sub>3</sub><sup>-</sup> in the aqueous phase is below 0.1, while for ratios in excess of 1 (also over eastern US), no enhancements are calculated (Fig. 4).

**Table 2.** Contributions of the different chemical processes to the simulated aerosol sulphate, nitrate and ammonium concentrations and their relative enhancement as a result of simulating realistic HONO levels during January in the eastern US and China.

| species                    | conditions                                    | eastern US, 95–65° W, 30–50° N  |                            |                | eastern China, 100–130° E, 25–55° N |                            |                |
|----------------------------|---|---------------------------------|----------------------------|----------------|-------------------------------------|----------------------------|----------------|
|                            |   | base_S [ $\mu\text{g m}^{-3}$ ] | S [ $\mu\text{g m}^{-3}$ ] | % <sup>a</sup> | base_S [ $\mu\text{g m}^{-3}$ ]     | S [ $\mu\text{g m}^{-3}$ ] | % <sup>a</sup> |
| total sulphate             | S1  | 2.97                            | 3.39                       | 14             | 5.18                                | 5.72                       | 11             |
|                            | S2  | 2.23                            | 2.92                       | 31             | 2.36                                | 3.17                       | 34             |
|                            | S3  | 2.47                            | 2.38                       | −3.7           | 4.26                                | 4.23                       | −0.61          |
|                            | H <sub>2</sub> SO <sub>4</sub> % <sup>b</sup> | 17                              | 30                         |                | 18                                  | 26                         |                |
|                            | Aqueous % <sup>c</sup>                        | 25                              | 14                         |                | 54                                  | 45                         |                |
| chemically formed sulphate | H <sub>2</sub> SO <sub>4</sub> %              | 41                              | 68                         |                | 25                                  | 37                         |                |
|                            | Aqueous %                                     | 59                              | 32                         |                | 75                                  | 63                         |                |
| nitrate                    | S1  | 2.05                            | 2.23                       | 8.9            | 2.51                                | 2.64                       | 4.9            |
|                            | S2  | 2.41                            | 2.61                       | 8.4            | 3.20                                | 3.25                       | 1.6            |
|                            | S3  | 1.79                            | 1.89                       | 5.9            | 2.40                                | 2.45                       | 2.4            |
| ammonium                   | S1  | 1.08                            | 1.25                       | 16             | 2.97                                | 3.18                       | 7.1            |
|                            | S2  | 0.77                            | 1.05                       | 37             | 1.56                                | 1.87                       | 20             |
|                            | S3  | 0.88                            | 0.86                       | −1.8           | 2.60                                | 2.53                       | −2.8           |
| number concentration       |   | base_S [ $\text{cm}^{-3}$ ]     | S [ $\text{cm}^{-3}$ ]     | % <sup>a</sup> | base_S [ $\text{cm}^{-3}$ ]         | S [ $\text{cm}^{-3}$ ]     | % <sup>a</sup> |
|                            | S1  | 2032                            | 2242                       | 10             | 5520                                | 6129                       | 11             |
|                            | S2  | 2129                            | 2675                       | 26             | 5830                                | 9857                       | 69             |
|                            | S3  | 1956                            | 1951                       | −0.26          | 5438                                | 5430                       | −0.15          |

S1: basic setup, all aerosol formation processes are included, S2: no aqueous phase chemistry, i.e., only uptake of gaseous species based on Henry's law constant, S3: no gas phase production of H<sub>2</sub>SO<sub>4</sub> from the oxidation of SO<sub>2</sub> with OH.

<sup>a</sup> Relative enhancement as a result of simulating realistic HONO levels.

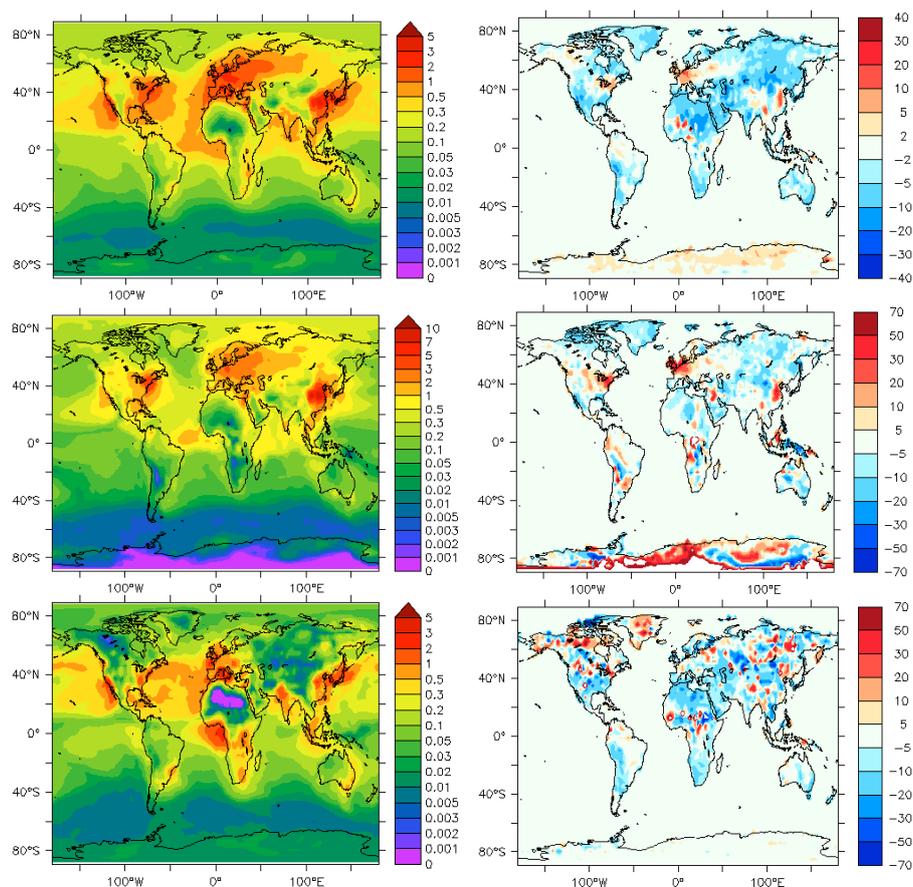
<sup>b</sup> Relative contribution of the gas phase H<sub>2</sub>SO<sub>4</sub>, calculated from the difference between S1 and S3.

<sup>c</sup> Relative contribution of the aqueous phase chemistry, calculated from the difference between S1 and S2.

In Table 3, the mean simulated concentrations of sulphate, nitrate and ammonium as well as their relative enhancements due to HONO are shown for the different soluble aerosol modes during summer and winter over the eastern US and China. The corresponding figures for the accumulation and coarse modes during winter are shown in Figs. 5 and 6, respectively. The relative enhancements in the aerosol mass concentrations are dominated by the nucleation and Aitken modes while the absolute enhancements are dominated by the accumulation mode and to a lesser extent the coarse mode. Since almost no enhancement in aerosol sulphate is calculated in the absence of the gas phase oxidation of SO<sub>2</sub> (Table 2, scenario S3), the enhancement in the larger modes is mainly related to the condensation of the enhanced gas phase H<sub>2</sub>SO<sub>4</sub> as well as the coagulation of smaller particles in the nucleation and Aitken modes. Sulphate aerosol concentrations have been reported to be underestimated by models, especially during winter with much better correlation between measured and simulated sulphates during summer (Jeuken, 2000; Pringle et al., 2010; Pozzer et al., 2012; de Meij et al., 2012). This underestimation has been related to a possible lack of photochemical oxidation pathways of SO<sub>2</sub>, which, based on our results, could be due to the underestimated OH during winter owing to the lack of HONO in their simulations (see Sect. 4.2).

#### 4.1.2 Nitrates

In coastal regions nitrate aerosols are predominantly found as sodium nitrate and more inland (continental) as ammonium nitrate (if there is an excess of ammonia). Nitrates are generally overestimated by models compared to reported measurements, partly owing to experimental artefacts related to nitrate sampling on quartz filters, especially in spring and summer. Nevertheless, better correlations with measurements were observed in summer compared to winter, which were partly explained by the absence of oxidation pathways (de Meij et al., 2012; Pozzer et al., 2012). In contrast to sulphates, which are produced mainly by the oxidation of S(IV) with H<sub>2</sub>O<sub>2</sub> and O<sub>3</sub> in clouds, nitrates are produced mainly from the transfer of gas phase nitric acid and N<sub>2</sub>O<sub>5</sub> to the clouds/particles. The impact of enhanced gas phase HNO<sub>3</sub> levels on aerosol nitrate depends on the availability of ammonia/ammonium, which is preferentially bound by H<sub>2</sub>SO<sub>4</sub>/sulphate (e.g., Seinfeld and Panids, 2006). In the absence of gas phase H<sub>2</sub>SO<sub>4</sub>, nitrate concentrations are reduced by about 13 and 5 % over eastern US and China, respectively, in the base run, which is related to the reduced surface area of sulphate aerosols allowing less condensation and hydrolysis (Table 2, S3 compared to S1).



**Fig. 7.** Simulated aerosol nitrate concentration ( $\mu\text{g m}^{-3}$ ) near the surface for the mean annual (upper panels), boreal winter (January, middle panels) and summer (July, lower panels) conditions and relative changes (in %, right panels) due to HONO enhancement in the model.

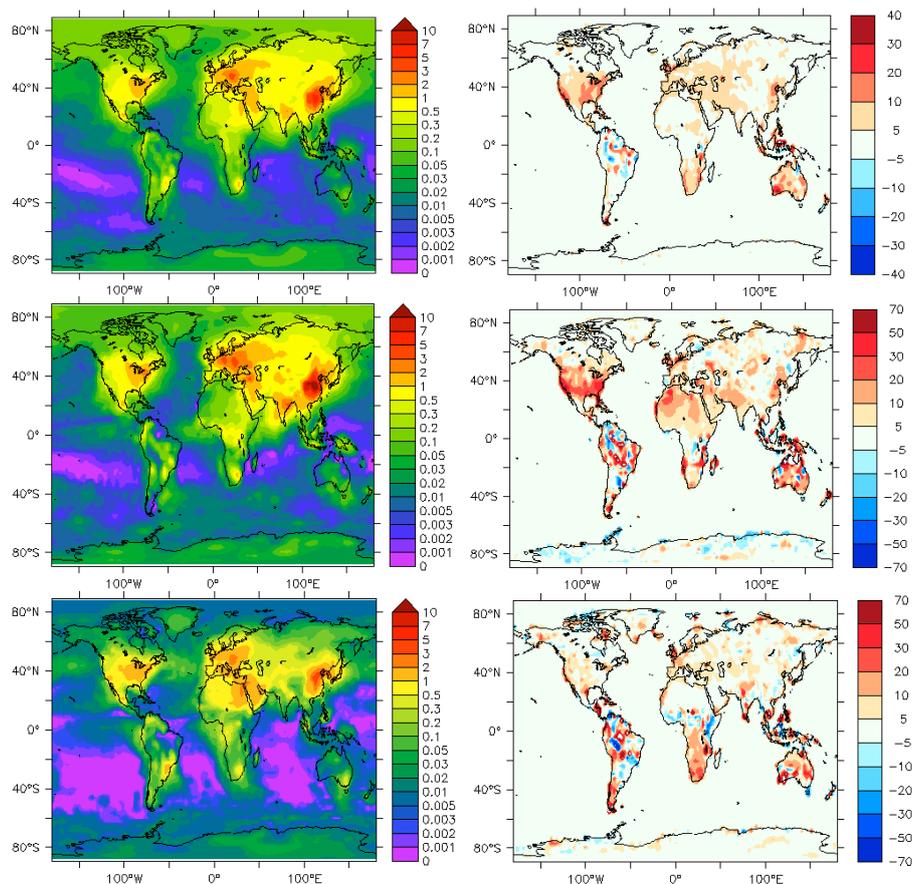
**Table 3.** Contributions of the different aerosol modes to the simulated aerosol mean sulphate, nitrate and ammonium concentrations and their relative enhancement as a result of simulating realistic HONO levels.

| aerosol mode  | date    | sulphate                |      |                            |      | nitrate                 |      |                            |        | ammonium                |      |                            |       |
|---------------|---------|-------------------------|------|----------------------------|------|-------------------------|------|----------------------------|--------|-------------------------|------|----------------------------|-------|
|               |         | eastern US <sup>1</sup> |      | eastern China <sup>2</sup> |      | eastern US <sup>1</sup> |      | eastern China <sup>2</sup> |        | eastern US <sup>1</sup> |      | eastern China <sup>2</sup> |       |
|               |         | $\mu\text{g m}^{-3}$    | %*   | $\mu\text{g m}^{-3}$       | %*   | $\mu\text{g m}^{-3}$    | %*   | $\mu\text{g m}^{-3}$       | %*     | $\mu\text{g m}^{-3}$    | %*   | $\mu\text{g m}^{-3}$       | %*    |
| total aerosol | January | 2.97                    | 14   | 5.18                       | 11   | 2.05                    | 9    | 2.51                       | 5      | 1.07                    | 16   | 2.97                       | 7     |
|               | July    | 3.07                    | 2    | 3.70                       | 1    | 0.49                    | -7   | 0.45                       | -8     | 0.89                    | 3    | 1.18                       | 1     |
| ns            | January | $6.13 \times 10^{-7}$   | 1021 | $4.74 \times 10^{-7}$      | 9090 | $8.79 \times 10^{-13}$  | 4564 | $1.56 \times 10^{-11}$     | 528353 | $1.90 \times 10^{-07}$  | 1098 | $1.59 \times 10^{-07}$     | 10303 |
|               | July    | $7.04 \times 10^{-9}$   | -20  | $4.31 \times 10^{-8}$      | 7    | $2.85 \times 10^{-18}$  | -85  | $3.88 \times 10^{-15}$     | -100   | $1.67 \times 10^{-09}$  | -24  | $1.51 \times 10^{-08}$     | 6     |
| ks            | January | $1.07 \times 10^{-2}$   | 65   | $1.47 \times 10^{-2}$      | 29   | $3.94 \times 10^{-06}$  | 134  | $1.76 \times 10^{-05}$     | 36     | $3.59 \times 10^{-03}$  | 75   | $4.79 \times 10^{-03}$     | 35    |
|               | July    | $1.24 \times 10^{-2}$   | 1    | $2.03 \times 10^{-2}$      | 4    | $2.46 \times 10^{-12}$  | -29  | $3.43 \times 10^{-11}$     | 18     | $4.42 \times 10^{-03}$  | 3    | $6.41 \times 10^{-03}$     | 4     |
| as            | January | 1.17                    | 31   | 3.27                       | 18   | 1.08                    | 23   | 1.58                       | 13     | 0.75                    | 25   | 2.13                       | 14    |
|               | July    | 2.22                    | 2    | 2.61                       | 0.9  | 0.26                    | -5   | 0.26                       | -10    | 0.79                    | 3    | 0.99                       | 1     |
| cs            | January | 1.79                    | 2    | 1.89                       | -3   | 0.97                    | -7   | 0.93                       | -8     | 0.32                    | -6   | 0.84                       | -10   |
|               | July    | 0.83                    | 1    | 1.07                       | 2    | 0.23                    | -9   | 0.19                       | -5     | 0.09                    | 3    | 0.18                       | 1     |

\* Relative enhancement as a result of simulating realistic HONO levels. Abbreviations "ns, ks, as and cs" refer to nucleation, Aitken, accumulation and coarse soluble aerosol modes, respectively. Data averaged over the eastern (1) US (95–65° W, 30–50° N) and (2) China (100–130° E, 25–55° N).

As shown in Figs. 5–7, nitrate aerosols are enhanced only in polluted continental regions (ammonia-rich regions), while only marginally affected or even decreased in all other areas (ammonia-limited environments). Similar to aerosol sulphate, the enhancement in nitrate aerosol mass is dom-

inated by the soluble accumulation and coarse modes (Table 3). Since most of the aerosol nitrate in the coastal and marine regions (ammonia-limited) is in the coarse mode, nitrate concentrations in this mode are suppressed owing to the enhancement in the sulphate concentrations



**Fig. 8.** Simulated aerosol ammonium concentration ( $\mu\text{g m}^{-3}$ ) near the surface for the mean annual (upper panels), boreal winter (January, middle panels) and summer (July, lower panels) conditions and relative changes (in %, right panels) due to HONO enhancement in the model.

while in the accumulation mode (mostly in continental polluted ammonia-rich environments) they are enhanced. Since  $\text{HNO}_3$  is most strongly enhanced during winter, nitrate is also mostly enhanced in this season while during summer only relatively small effects are simulated.

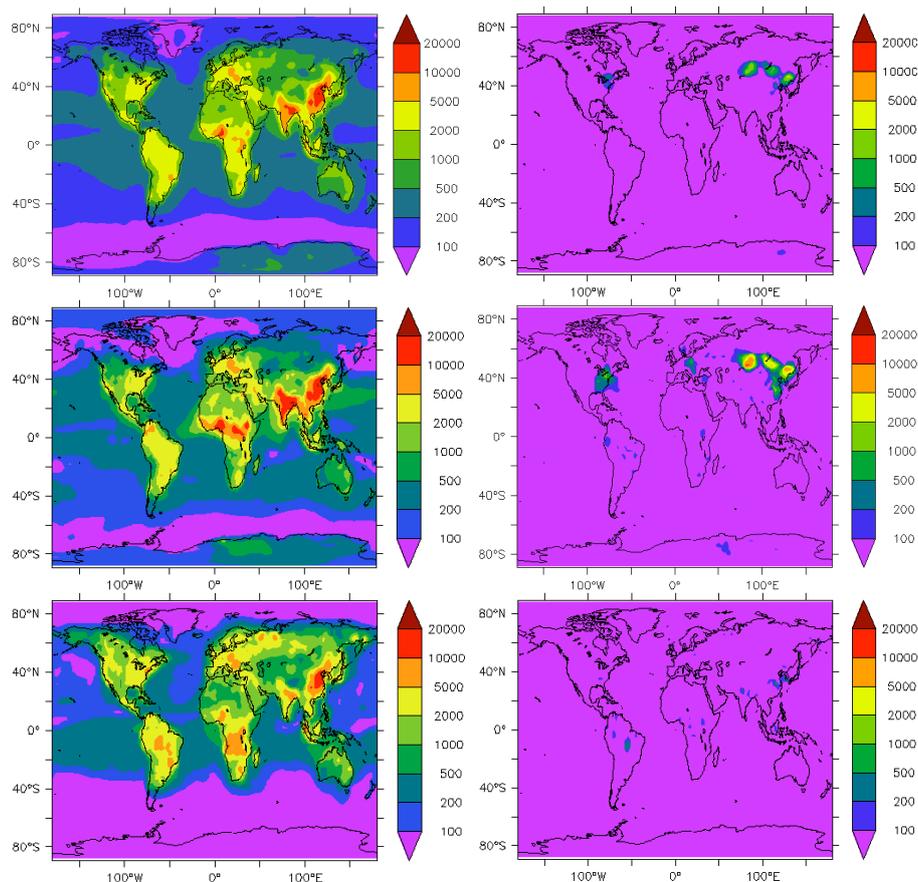
#### 4.1.3 Ammonium

Ammonium aerosols are largely restricted to the continental regions, with the highest concentrations in the US, Europe, China and India (Fig. 8), mostly due to ammonia ( $\text{NH}_3$ ) emissions from agriculture. Ammonium aerosols form preferably as ammonium sulphate but also as ammonium nitrate if there is an excess of cations (Pringle et al., 2010). Ammonium aerosol can be enhanced as a result of additional sulphate and nitrate (in high  $\text{NO}_x$  regions) only in ammonia-rich regions while in ammonia-limited regions no significant effects are expected. As shown in Fig. 8, the total load of ammonium aerosols is enhanced only in high  $\text{NO}_x$ , ammonia-rich environments, i.e., mainly in the US, western Europe and eastern Asia. Though ammonia reaches its global maximum levels in India, only small relative ammonium en-

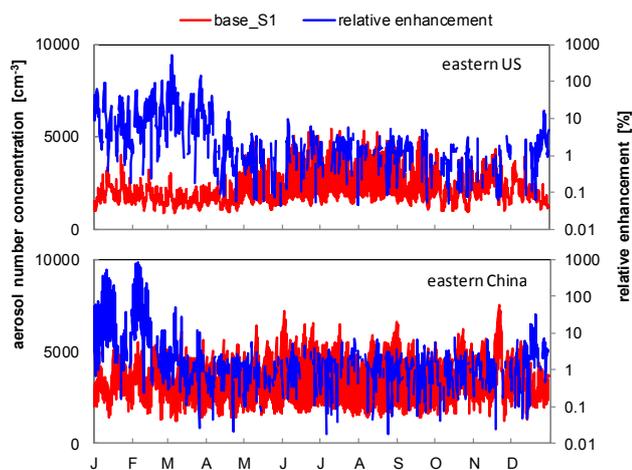
hancement occurs in response to HONO induced oxidation due to the relatively low  $\text{NO}$  levels in this region (see Sect. 4 and e.g., Ojha et al., 2012). The enhancement in ammonium reaches 20 % annually and up to 50 % in winter while in summer the effects are negligible.

#### 4.1.4 Aerosol number concentrations

As shown in Fig. 9, the mean near-surface aerosol number concentrations are also significantly enhanced; most apparently in high  $\text{NO}_x$  polluted regions with an annual mean relative enhancement over the eastern US and eastern Asia by up to about 70 %. Furthermore, the highest relative changes in aerosol number concentrations are simulated during winter months, reaching about 100 % over eastern Asia while during summer almost no enhancements are calculated. The simulated seasonal cycle of aerosol number concentrations in the base run as well as the relative enhancements from simulating realistic HONO over the eastern US and China are shown in Fig. 10, and the absolute differences during summer and winter are shown in Fig. 11. The maximum relative enhancements are simulated in both regions during winter



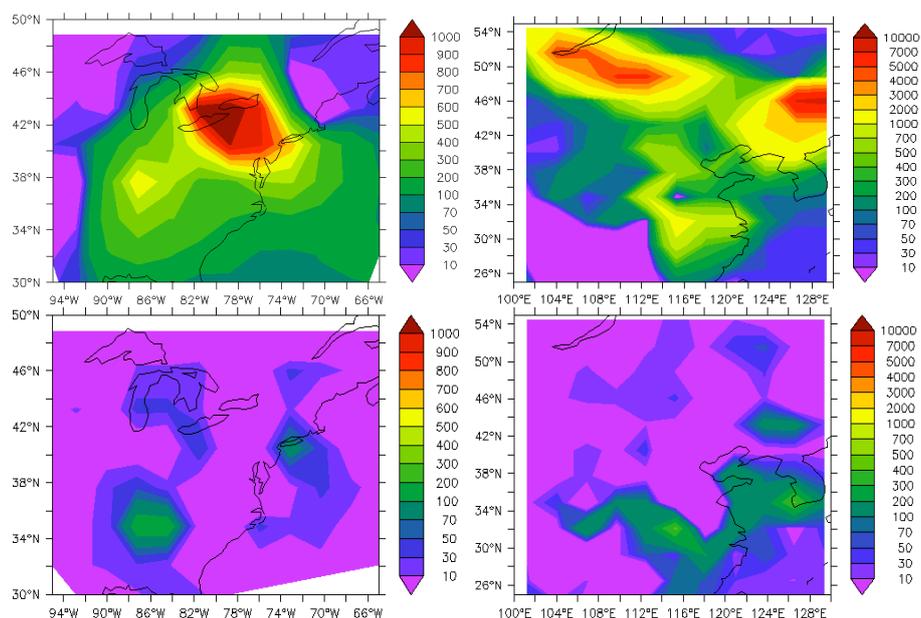
**Fig. 9.** Simulated mean aerosol number concentrations ( $\text{cm}^{-3}$ ) near the surface for the mean annual (upper panels), boreal winter (January, middle panels) and summer (July, lower panels) conditions and the absolute changes (right panels) due to HONO enhancement in the model.



**Fig. 10.** Simulated aerosol number concentrations and the relative enhancement related to simulating realistic HONO levels near the surface over the eastern US ( $95\text{--}65^\circ\text{W}$  and  $30\text{--}50^\circ\text{N}$ , upper panels) and China ( $100\text{--}130^\circ\text{E}$  and  $25\text{--}55^\circ\text{N}$ , lower panels).

while during summer the enhancements are again negligible. The increase of aerosol number concentrations over eastern China reaches about  $10^4\text{ cm}^{-3}$  during January, about one order of magnitude higher than that over the eastern US while during summer enhancements in both regions do not exceed  $500\text{ cm}^{-3}$  (Fig. 11).

The mean simulated aerosol number concentration for the different aerosol modes over the eastern US and China are shown in Table 4. Aerosol number concentrations are dominated by the Aitken and accumulation modes, reflecting typical distributions of the near-surface aerosol number concentrations (see e.g., Pringle et al., 2010, and references therein). The mean relative enhancement in the total aerosol number concentrations over the eastern US and China is about 10–11 % (see Table 4). Another important effect is the transformation from hydrophobic to hydrophilic aerosols related mainly to the enhanced condensation of the gas phase constituents (e.g.,  $\text{H}_2\text{SO}_4$ ,  $\text{HNO}_3$ ) as well as the coagulation of the soluble new aerosol particles with the hydrophobic aerosol particles followed by transfer into the hydrophilic mode. For example, over eastern China the particle number concentration in the hydrophobic Aitken (ki) mode is



**Fig. 11.** Simulated monthly mean change in the aerosol number concentrations ( $\text{cm}^{-3}$ ) near the surface over the eastern US (left) and China (right) during winter (January, upper panels) and summer (July, lower panels) due to HONO enhancement in the model.

reduced by about 6 % (about  $245 \text{ cm}^{-3}$ ) while that of the hydrophilic ks and as modes are enhanced by 31 and 14 %, respectively (total about  $334 \text{ cm}^{-3}$ ).

Similarly, over the eastern US the hydrophobic Aitken mode particle number is reduced by 13 % (about  $161 \text{ cm}^{-3}$ ) while the hydrophilic ks and as modes are enhanced by 44 and 27 %, respectively (total about  $306 \text{ cm}^{-3}$ ), see Table 4. In the absence of gas phase  $\text{H}_2\text{SO}_4$ , almost no enhancements in the total aerosol number concentrations are calculated as a result of realistic HONO levels. The total enhancement in the aerosol number concentrations can thus be attributed to the strong enhancement in the nucleation mode, especially over eastern China, related mainly to the enhancement in the gas phase sulphuric acid concentration (i.e., nucleation of new particles is calculated in the model as a function of the temperature, relative humidity and the concentration of gas phase  $\text{H}_2\text{SO}_4$ , see Pringle et al., 2010).

The enhancement of the aerosol number and solubility implies potential impacts on the cloud nucleation properties and particle lifetime. The enhancement of fine mode aerosol (nucleation, Aitken and accumulation mode) particles, as a result of simulating realistic HONO levels, may directly affect climate by increasing the light scattering and indirectly via altering cloud properties.

#### 4.2 Comparison with large scale measurements

A summary of the comparisons between the mean simulated aerosol sulphates, nitrates and ammonium by the base\_S1 and S1 model runs and the measurements by CASTNET, EMEP and EANET networks during winter (January to

**Table 4.** Contributions of the different aerosol modes to the simulated mean aerosol number concentrations and their relative enhancement as a result of simulating realistic HONO levels during January.

| aerosol mode  | run | aerosol number concentration |      |                            |      |
|---------------|-----|------------------------------|------|----------------------------|------|
|               |     | eastern US <sup>1</sup>      |      | eastern China <sup>2</sup> |      |
|               |     | $\text{cm}^{-3}$             | %*   | $\text{cm}^{-3}$           | %*   |
| total aerosol | S1  | 2032                         | 10   | 5520                       | 11   |
|               | S3  | 1956                         | −0.3 | 5438                       | −0.1 |
| ns            | S1  | 8.51                         | 681  | 10.4                       | 5131 |
|               | S3  | $7.59 \times 10^{-4}$        | 48   | $1.07 \times 10^{-3}$      | −37  |
| ks            | S1  | 552                          | 44   | 825                        | 31   |
|               | S3  | 358                          | 2    | 550                        | 1    |
| as            | S1  | 235                          | 27   | 567                        | 14   |
|               | S3  | 148                          | 1    | 369                        | 2    |
| cs            | S1  | 1.40                         | 0.6  | 2.61                       | −2   |
|               | S3  | 1.33                         | 0.2  | 2.70                       | −4   |
| ki            | S1  | 1235                         | −13  | 4075                       | −6   |
|               | S3  | 1448                         | −1   | 4472                       | −1   |
| ai            | S1  | $5.97 \times 10^{-3}$        | −38  | 33.9                       | 2    |
|               | S3  | $1.15 \times 10^{-1}$        | −2   | 39.1                       | 2    |
| ci            | S1  | $1.08 \times 10^{-3}$        | −30  | 2.74                       | 2    |
|               | S3  | $5.73 \times 10^{-3}$        | −1   | 3.10                       | 1    |

\* Relative enhancement as a result of simulating realistic HONO levels (S1 run). Abbreviations “ns, ks, as and cs” refer to nucleation, Aitken, accumulation and coarse soluble aerosol modes, respectively, while “ki, ai and ci” refer to the corresponding insoluble modes. Data are averaged over the eastern (1) US (95–65° W, 30–50° N) and (2) China (100–130° E, 25–55° N).

**Table 5.** Summary of the comparison of the base\_S1 and S1 model simulations to observations of aerosol concentrations during January–March 2001. OAM and MAM are the arithmetic mean of the observations and of the model, respectively, in  $\mu\text{g m}^{-3}$ , while OSTD and MSTD are the standard deviation of the observations and model, respectively, in  $\mu\text{g m}^{-3}$ . MAM, OAM, MSTD and OSTD represent co-located measurements and model results (i.e., based on the locations of the observations). PF2 is the percentage of modelled point within a factor of two of the observations. RMS denotes the Root-Mean-Square error.

| species            | network | Nr. of stations | OAM  | OSTD | MAM [ $\mu\text{g m}^{-3}$ ] |      | MSTD    |      | MAM/OAM |      | PF2 [%] |    | RMS     |      |
|--------------------|---------|-----------------|------|------|------------------------------|------|---------|------|---------|------|---------|----|---------|------|
|                    |         |                 |      |      | base_S1                      | S1   | base_S1 | S1   | base_S1 | S1   | base_S1 | S1 | base_S1 | S1   |
| $\text{SO}_4^{2-}$ | CASTNET | 53              | 3.17 | 2.17 | 2.57                         | 3.05 | 0.75    | 0.93 | 0.81    | 0.96 | 88      | 85 | 1.92    | 1.82 |
|                    | EMEP    | 80              | 2.00 | 1.18 | 2.95                         | 3.09 | 1.29    | 1.41 | 1.48    | 1.55 | 73      | 71 | 1.27    | 1.32 |
|                    | EANET   | 11              | 2.82 | 1.73 | 2.89                         | 3.13 | 1.70    | 1.96 | 1.03    | 1.11 | 80      | 77 | 2.12    | 2.29 |
| $\text{NO}_3^-$    | CASTNET | 53              | 1.27 | 1.53 | 1.81                         | 1.97 | 1.07    | 1.26 | 1.43    | 1.55 | 45      | 45 | 1.71    | 1.81 |
|                    | EMEP    | 27              | 1.75 | 2.16 | 2.24                         | 2.42 | 1.00    | 1.19 | 1.28    | 1.38 | 40      | 39 | 2.00    | 1.96 |
|                    | EANET   | 10              | 0.81 | 0.99 | 1.26                         | 1.14 | 1.05    | 0.95 | 1.55    | 1.41 | 63      | 63 | 1.68    | 1.59 |
| $\text{NH}_4^+$    | CASTNET | 53              | 1.26 | 0.83 | 1.23                         | 1.43 | 0.66    | 0.76 | 0.98    | 1.14 | 86      | 80 | 0.70    | 0.70 |
|                    | EMEP    | 21              | 1.00 | 0.99 | 1.43                         | 1.54 | 0.92    | 1.01 | 1.43    | 1.54 | 54      | 51 | 0.87    | 0.90 |
|                    | EANET   | 9               | 0.96 | 0.71 | 0.96                         | 1.02 | 0.86    | 0.93 | 1.01    | 1.07 | 66      | 66 | 0.89    | 0.94 |

March) 2001 are listed in Table 5. As shown in Table 5, simulating realistic HONO levels (S1 simulations) generally increases the mean modelled/measured (MAM/OAM) ratio, most significantly over the US (about 20 %), thus improving the agreement between modelled and measured sulphate concentrations in this region (i.e., modelled to observed ratio changes from 0.81 to 0.96). This is due to the enhanced gas phase oxidation of  $\text{SO}_2$ , being most important over the eastern US, i.e., compared to eastern China (see Sect. 4.1.1 and Table 2). Simulated nitrate by both the base\_S1 and S1 runs are generally overestimated with model/measured ratio ranges from 1.28 to 1.55. These overestimations of nitrate are however related to known issues, including measurement biases, overestimation of sea spray aerosols by GMXe, which may result in overestimation of aerosol nitrate by the uptake of  $\text{HNO}_3$  (Pringle et al., 2010; de Meij et al., 2012). Over eastern Asia (EANET), the modelled/observed ratio for nitrate decreases from 1.55 for the base\_S1 to 1.41 for the S1 run (i.e., as a result of simulating realistic HONO levels). Ammonium aerosols are well simulated by the reference run over the US and eastern Asia (MAM/OAM are 0.98 and 1.01, respectively) while slightly overestimated over Europe, in agreement with Pringle et al. (2010). The ratio of simulated/measured ammonium aerosol generally follows that of sulphate and increases in the S1 run compared to the reference base\_S1, especially over the US. Pozzer et al. (2012) investigated the EMAC simulation results for the years 2005–2008 and showed also that the model tends to underestimate sulphate aerosol compared to CASTNET and also compared to EMEP and EANET measurements (Pozzer et al., 2012). Furthermore, Pringle et al. (2010) compared simulated concentrations of sulphates, nitrates and ammonium fine mode aerosol ( $< 1 \mu\text{m}$  diameter) concentrations with aerosol mass spectrometer data and found that the model tends to underestimate the three measured species with sulphates being most poorly captured. Since the HONO enhancements are

most significant for the accumulation mode (see Table 3), it can be expected that with realistic HONO simulations, these simulated concentrations are in better agreement with the mass spectrometer measurements. They showed also that the EMAC model tends to overestimate nitrate in coastal regions, which (see Sect. 4.1.2) is more realistically simulated by improving the HONO simulation. Therefore, by accounting for realistic HONO levels the simulated aerosol concentrations are improved compared to measurements.

## 5 Conclusions

In this study the impact of our realistic HONO representation on the simulated aerosol composition has been investigated. By applying our parameterization, simulated HONO mixing ratios typically increase by an order of magnitude, in agreement with our previous study, which showed that this parameterization leads to much better agreement with measurements compared to the model run that only considers the reaction of  $\text{NO} + \text{OH}$  as a source of HONO. Owing to the enhanced gas phase HONO photolysis,  $\text{HO}_x$  ( $\text{OH} + \text{HO}_2$ ) levels are enhanced, thus promoting the formation of organic (e.g., PAN, OVOC) and inorganic ( $\text{H}_2\text{O}_2$ ,  $\text{O}_3$ ,  $\text{HNO}_3$  and  $\text{H}_2\text{SO}_4$ ) oxidation products in polluted regions, especially in winter when other photolytic  $\text{HO}_x$  sources are of relatively minor importance.

Simulating realistic HONO levels is found to significantly enhance the near surface aerosol sulphate concentrations, mainly due to enhanced gas phase oxidation of  $\text{SO}_2$  with OH. Nitrate and ammonium are also significantly enhanced in ammonia-rich regions, while in ammonia-limited regions the effect is small. Furthermore, aerosol number concentrations are also significantly higher, accompanied by transfer from the hydrophobic to hydrophilic modes, which is mainly related to the condensation of  $\text{H}_2\text{SO}_4$  on the hydrophobic particles. The enhancement of the aerosol number and solubility

implies potential impacts on the cloud nucleating properties and the particle lifetime.

The relative enhancements of all species reach a maximum during the winter season while generally being negligible during summer. The simulation results are compared to data from the monitoring networks CASTNET, EMEP and EANET. The model results are in reasonable agreement with measurements and in line with previous evaluations. Simulating realistic HONO levels was found to enhance the species  $\text{SO}_4^{2-}$ ,  $\text{NO}_3^-$  and  $\text{NH}_4^+$ , with sulphates being most strongly enhanced, especially over the US, improving the agreement with measurements in this region. The results of the study underscore the importance of HONO for the atmospheric oxidizing capacity and the central role of cloud-aerosol interactions in aerosol formation and growth.

**Supplementary material related to this article is available online at <http://www.atmos-chem-phys.net/14/1167/2014/acp-14-1167-2014-supplement.pdf>.**

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