



## Secondary organic aerosol formation from biomass burning intermediates: phenol and methoxyphenols

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

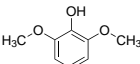
**Abstract.** The formation of secondary organic aerosol from oxidation of phenol, guaiacol (2-methoxyphenol), and syringol (2,6-dimethoxyphenol), major components of biomass burning, is described. Photooxidation experiments were conducted in the Caltech laboratory chambers under low-NO<sub>x</sub> (< 10 ppb) conditions using H<sub>2</sub>O<sub>2</sub> as the OH source. Secondary organic aerosol (SOA) yields (ratio of mass of SOA formed to mass of primary organic reacted) greater than 25 % are observed. Aerosol growth is rapid and linear with the primary organic conversion, consistent with the formation of essentially non-volatile products. Gas- and aerosol-phase oxidation products from the guaiacol system provide insight into the chemical mechanisms responsible for SOA formation. Syringol SOA yields are lower than those of phenol and guaiacol, likely due to novel methoxy group chemistry that leads to early fragmentation in the gas-phase photooxidation. Atomic oxygen to carbon (O:C) ratios calculated from high-resolution-time-of-flight Aerodyne Aerosol Mass Spectrometer (HR-ToF-AMS) measurements of the SOA in all three systems are ~ 0.9, which represent among the highest such ratios achieved in laboratory chamber experiments and are similar to that of aged atmospheric organic aerosol. The global contribution of SOA from intermediate volatility and semivolatile organic compounds has been shown to be substantial (Pye and Seinfeld, 2010). An approach to rep-

resenting SOA formation from biomass burning emissions in atmospheric models could involve one or more surrogate species for which aerosol formation under well-controlled conditions has been quantified. The present work provides data for such an approach.

### 1 Introduction

Biomass burning is a major source of atmospheric organic aerosol (OA), with contributions from both anthropogenic (biofuel, deforestation, etc.) as well as natural sources such as wildfires. Aerosol produced from biomass burning has been estimated to account for 90 % of all primary organic carbon emitted globally from combustion sources (Ito and Penner, 2005). Bond et al. (2004) estimated global annual emissions of carbonaceous aerosol due to combustion as 8.0 Tg for black carbon (BC) and 33.9 Tg organic carbon (OC). Biofuel and open burning were estimated by these authors to account, respectively, for 20 % and 42 % of BC, and 19 % and 74 %, of OC. The amount and composition of organic aerosol formed from these sources vary widely based on combustion conditions and fuel type. Measured wood smoke emissions from fireplaces, for example, depend strongly on the type of fuel consumed and can vary in magnitude by a factor of five for

Table 1. Chemical properties.

Compound	Structure	Boiling Pt. <sup>a</sup> (°C)	V.P. <sup>a</sup> @ 25 °C (mm Hg)	$k_{\text{OH}} \times 10^{11}$ (molec cm <sup>-3</sup> s <sup>-1</sup> )
Phenol		182	0.351	2.7 ± 0.25 at 25 °C (Calvert et al., 2002)
Guaiacol		205	0.13	7.53 ± 0.41 at 21 ± 2 °C (Coeur-Tourneur et al., 2010)
Syringol		261	0.00602 <sup>b</sup>	9.66 ± 1.11 at 21 ± 2 °C (Lauraguais et al., 2012)

<sup>a</sup> Experimental database values within EPA (2012)

<sup>b</sup> Calculated within EPA (2012)

different fuel sources and sampling techniques (Fine et al., 2001). Particulate matter from biomass burning, including wildfires, residential wood burning, and deforestation is also of major concern for visibility, climate change and health effects (Naeher et al., 2007).

Recent studies have been devoted to measuring and chemically characterizing the gas and particle-phase emissions from biomass burning in-situ (Iinuma et al., 2010; Akagi et al., 2012) and in the laboratory (Veres et al., 2010; Iinuma et al., 2010). Understanding the fate of such emissions in the atmosphere as a plume rises and advects away from the source is a challenge to discern. One approach towards addressing this problem is to simulate the photooxidation of known biomass burning tracers in the laboratory under highly-controlled chemical conditions.

Phenols are produced from the pyrolysis of lignin, the second most abundant polymer on Earth comprising the secondary cell wall of plants. Hawthorne et al. (1989) showed that phenols and methoxyphenols account for a significant fraction of particulate matter derived from wood smoke, constituting 21 % and 45 % by mass of total aerosol, respectively. Emission rates for methoxyphenol species from biomass burning have been estimated to range from 420–900 mg kg<sup>-1</sup> fuel (Hawthorne et al., 1989; Hawthorne et al., 1992; Schauer et al., 2001). In addition to the significant emission factors of these compounds, guaiacol, which is emitted from both hard and soft wood combustion, and syringol, which is emitted primarily from soft woods, have been proposed as aerosol markers for wood combustion as well as biomarkers to determine human exposure levels (Dills et al., 2001; Simpson and Naeher, 2010).

Such phenolic compounds are used in ozonolysis and photooxidation studies to simulate the chemical fate and aerosol forming potential of such biomass burning intermediates in the atmosphere. Often these include candidate compounds such as phenol and catechol (Atkinson et al., 1992; Olariu, 2001; Olariu et al., 2002; Berndt and Boge, 2003; Tomas et al., 2003; Coeur-Tourneur et al., 2009; Iinuma et al., 2010; Nakao et al., 2011; Ofner et al., 2011). Many of these com-

pounds can also serve as model systems for understanding aerosol formation from humic substances, the role of water, and associated heterogenous processes (Ammann et al., 2005; Nieto-Gligorovski et al., 2008; Ofner et al., 2010, 2011; Sun et al., 2010).

Still, little is known about the detailed gas-phase photooxidation mechanism of methoxyphenol compounds. Here we investigate the aerosol-forming potential of phenol and two of the most abundant methoxyphenol species, guaiacol and syringol (Table 1). Both guaiacol and phenol are intermediate volatility organic compounds; syringol has a vapor pressure 100 times less than those of phenol and guaiacol. The molecular structures are shown in Table 1 along with vapor pressures and reaction rate constants with OH.

We determine the fractional yield of secondary organic aerosol (SOA) from photooxidation of phenol, guaiacol, and syringol by OH in laboratory chamber studies under low-NO<sub>x</sub> ([NO<sub>x</sub>]<sub>0</sub> < 10 ppb) conditions. While NO<sub>x</sub> is certainly present in biomass burning events at higher levels than this, we focus on the results of the low-NO<sub>x</sub> experiments in the main text and leave discussion on the effect of NO<sub>x</sub> in the Appendix. The atmospheric chamber experiments performed on these systems were performed under idealized chemical conditions such that specific regimes of chemistry could be isolated. That is, low-NO<sub>x</sub> experiments resulted in the predominant fate of RO<sub>2</sub> radical to be reaction with HO<sub>2</sub>, and in the high-NO<sub>x</sub> experiments, reaction with NO<sub>x</sub>. In the atmosphere there exists a spectrum of chemical regimes that depend on proximity to the heart of the biomass burning source. Owing to the high-NO<sub>x</sub> experiments resulting in levels of total NO<sub>x</sub> approaching 1 ppm (beyond typical atmospheric levels), we find the results to be chemically instructive, though not practical for use as an accurate representation of the SOA formation from these systems under atmospherically relevant NO<sub>x</sub> conditions. For example, Iinuma et al. (2010) measured co-located NO<sub>x</sub> to be on the order of 20 ppb.

Studies on the photooxidation mechanisms of aromatics and phenol in general find that the product distributions for aromatics chemistry are qualitatively similar in the absence

**Table 2.** Experimental conditions and aerosol yields.

Date (2010)	Organic	[Org] <sub>0</sub> (ppb)	[Org] <sub>f</sub> (ppb)	RH <sub>0</sub> (%)	[NO <sub>2</sub> ] <sub>0</sub> (ppb)	[NO] <sub>0</sub> (ppb)	Vol <sub>0</sub> <sup>e</sup> μm <sup>3</sup> cm <sup>-3</sup>	Vol <sub>f</sub> <sup>e</sup> μm <sup>3</sup> cm <sup>-3</sup>	Yield
1/29	guaiacol	66.3	13.4	5.6	< LDL <sup>d</sup>	< LDL <sup>d</sup>	18.2	120.0	0.49 ± 0.01
2/2	guaiacol	5.9	< LDL <sup>a</sup>	5.8	< LDL <sup>d</sup>	< LDL <sup>d</sup>	16.4	25.0	0.46 ± 0.03
2/4	guaiacol	12.4	< LDL <sup>a</sup>	5.5	< LDL <sup>d</sup>	< LDL <sup>d</sup>	13.6	31.0	0.44 ± 0.03
2/6	guaiacol	45.5	4.1	6.7	< LDL <sup>d</sup>	5	12.9	85.0	0.50 ± 0.02
2/11	phenol	47.6	9.9	5.7	< LDL <sup>d</sup>	6	16.0	58.0	0.44 ± 0.08
2/15	phenol	10.0	2.1	6.2	< LDL <sup>d</sup>	5	11.4	20.0	0.40 ± 0.06
2/17	phenol	73.9	14.4	< LDL <sup>c</sup>	< LDL <sup>d</sup>	< LDL <sup>d</sup>	12.0	32.9	0.25 ± 0.01
2/20	phenol	101.9	18.8	< LDL <sup>c</sup>	< LDL <sup>d</sup>	5	10.0	69.0	0.24 ± 0.03
3/10	syringol	185.1	< LDL <sup>b</sup>	< LDL <sup>c</sup>	7	< LDL <sup>d</sup>	11.0	325.0	0.37 ± 0.01
3/15	syringol	49.5	< LDL <sup>b</sup>	< LDL <sup>c</sup>	< LDL <sup>d</sup>	< LDL <sup>d</sup>	14.2	67.0	0.25 ± 0.01
3/29	syringol	112.9	< LDL <sup>b</sup>	< LDL <sup>c</sup>	9	< LDL <sup>d</sup>	11.7	174.0	0.34 ± 0.01

<sup>a</sup> below CIMS lower detection limit (LDL) of 0.1 ppb for guaiacol

<sup>b</sup> CIMS LDL = 0.5 ppb for syringol

<sup>c</sup> RH probe LDL = 5 % RH

<sup>d</sup> NO<sub>x</sub> Analyzer LDL = 5 ppb

<sup>e</sup> Vol<sub>0</sub> is initial seed volume; Vol<sub>f</sub> is final volume (seed + organic aerosol), as measured by DMA

and presence of NO<sub>x</sub> (Atkinson et al., 1992; Olariu et al., 2002; Sato et al., 2007), with noted presence of nitroaromatics generated in the presence of NO<sub>x</sub>. Still, these products tend to have gas-phase yields less than 0.1 (Olariu et al., 2002; Berndt and Boge, 2006). Thus, the low-NO<sub>x</sub> data presented here focus analyses on relevant gas- and particle-phase constituents in guaiacol photooxidation to elucidate the chemical mechanism involved in guaiacol SOA formation. This chemistry may best represent the chemical regime in which biomass burning emissions are transported to areas where NO<sub>x</sub> levels have become diluted and photochemistry ensues. We also qualitatively compare the chemistry of the tested compounds to explore why syringol results in lower SOA yield compared to phenol and guaiacol. Further motivation for this work is the potential use of phenol, guaiacol, and syringol as compounds to represent biomass burning emissions in atmospheric models of organic aerosol formation. SOA products identified in the laboratory studies may also serve as markers for biomass burning in ambient aerosols.

## 2 Experimental section

### 2.1 Chamber experiments

All experiments were carried out in the Caltech dual 28 m<sup>3</sup> Teflon chambers. Details of the facilities have been described previously (Cocker et al., 2001; Keywood et al., 2004). Before each experiment, the chambers were flushed with dried purified air RH < 5 %, NO<sub>x</sub> < 2 ppb for > 24 h, until the particle number concentration was < 100 cm<sup>-3</sup> and the volume concentration was < 0.1 μm<sup>3</sup> cm<sup>-3</sup>. In all yield experiments, ammonium sulfate seed aerosol was used to promote conden-

sation of low volatility oxidation products. The seed aerosol was generated by atomization of a 0.015 M aqueous ammonium sulfate solution. Following atomization, the size distribution of the seed particles peaked at ~ 56 nm with an average number concentration of ~ 2700 cm<sup>-3</sup>, and a total seed volume concentration of 10–15 μm<sup>3</sup> cm<sup>-3</sup> was achieved. The hydrocarbon was introduced into the chamber by injecting a known volume of pure hydrocarbon into a glass bulb and flowing purified air over the hydrocarbon at 5 L min<sup>-1</sup> until the hydrocarbon had fully vaporized. Gentle heat was applied to the glass bulb to support evaporation of syringol (solid) into the chamber.

For these low-NO<sub>x</sub> experiments, hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) was used as the OH precursor. We refer to the reaction conditions of these experiments as “low-NO<sub>x</sub>” because there is no intentional injection of NO<sub>x</sub> ([NO<sub>x</sub>]<sub>0</sub> < 5 ppb, the lower detection limit of the NO<sub>x</sub> analyzer). There are three hundred 40 W black lights, Sylvania 350BL, with an emission spectrum that peaks at 340–350 nm wavelength. Based on the irradiance spectrum of the lights and the absorption cross section of H<sub>2</sub>O<sub>2</sub> (Sander, 2011), this results in a photolysis rate constant,  $j_{\text{H}_2\text{O}_2} = 2.9 \times 10^6 \text{ s}^{-1}$ . Prior to atomization of the ammonium sulfate seed, H<sub>2</sub>O<sub>2</sub> was introduced by flowing 5 L min<sup>-1</sup> of purified air through a glass trap containing 280 μL of a 50 % aqueous H<sub>2</sub>O<sub>2</sub> solution. The glass trap was submerged in a warm water bath maintained at 35–38 °C. This resulted in an approximate concentration of 4 ppm H<sub>2</sub>O<sub>2</sub> in the chamber.

The aerosol number concentration and size distribution were measured by a differential mobility analyzer (DMA, TSI model 3081) coupled with a condensation nuclei counter (TSI, CNC-3760). After allowing all concentrations to

stabilize, irradiation was initiated. The temperature ( $T$ ), relative humidity (RH), and concentrations of  $O_3$ ,  $NO$ , and  $NO_x$  were continuously monitored. Experiments were run at temperatures ranging 20–26 °C and varied within  $\pm 2$  °C. RH remained below 10 %. The lower detection limits (LDL) of the utilized analyzers are 5 % RH, 2 ppb  $O_3$ , and 5 ppb  $NO_x$ . Table 2 summarizes the experimental conditions for the series of methoxyphenol oxidation experiments conducted.

## 2.2 Gas-phase measurements

### 2.2.1 Gas Chromatography/Flame-Ionization Detection (GC/FID)

The hydrocarbon concentration was continuously monitored by GC/FID in the phenol experiments. Chamber air was sampled into a 10 mL injection loop and injected onto a HP5 15 m  $\times$  0.53 mm ID  $\times$  1  $\mu$ m thickness column installed on a 6890 Agilent GC. The GC response was calibrated by dissolving a known mass of the hydrocarbon in methanol and then vaporizing a known volume of that solution into a 55 L Teflon chamber. Guaiacol and syringol measurements obtained using the GC/FID were unreliable due to condensation loss in the sample loop; thus, the hydrocarbon concentration during these experiments was monitored using Chemical Ionization Mass Spectrometry (CIMS) in negative mode operation. The LDL for phenol on the GC-FID was determined to be  $\sim 1$  ppb from calibrations.

### 2.2.2 Chemical Ionization Mass Spectrometry (CIMS)

Monitoring of gas-phase oxidation products was carried out in real time by the use of a CIMS instrument. The details of this instrument are described elsewhere (St. Clair et al., 2010). The instrument operates in both negative mode, using  $CF_3O^-$  as a reagent ion, and in positive proton transfer reaction (PTR)-MS mode. Negative mode is found to be more selective towards the detection of hydroperoxides and polar molecules, particularly acids, while positive mode detects a broader range of organic compounds. Analytes in negative mode can be monitored as the cluster product  $[R \cdot CF_3O]^-$  and/or as the transfer product if it is more strongly acidic  $[R \cdot F]^-$ , where R is the analyte. Analytes in positive mode cluster as  $[R \cdot (H_2O)_n]^+$ . Mass scans covering masses 50–300 amu for negative mode, and 50–200 amu for positive mode, with a total scan time of  $\sim 6$  min were continuously repeated over the course of each experiment.

Guaiacol was monitored at the fluoride transfer product ( $[M + 19]^-$ ),  $m/z$  143, and the cluster product ( $[M + 85]^-$ ),  $m/z$  209, in negative mode operation. Syringol was also monitored at the fluoride transfer product,  $m/z$  173, and the cluster product,  $m/z$  239. Calibrations of guaiacol and syringol on the CIMS were performed similar to those of phenol on the GC-FID by preparing 55 L Teflon chamber standards of varying concentrations of the methoxyphenol compounds. The

LDL for guaiacol on the CIMS was established to be 100 ppt and that for syringol to be 500 ppt. While phenol was detected at both the transfer ( $m/z$  113) and the cluster product ( $m/z$  179), a calibration curve was not made because a photooxidation product in the phenol system interfered significantly at  $m/z$  113.

## 2.3 Particle-phase measurements

### 2.3.1 Chamber filter sample collection, extraction, and off-line chemical characterization

A detailed description of the aerosol filter sample collection and extraction protocol for the Caltech laboratory chambers has been previously published (Surratt et al., 2008). Aerosol samples were collected on Teflon filters (PALL Life Sciences, 47 mm diameter, 1.0  $\mu$ m pore size, teflomembrane). Filter samplers employed for aerosol filter sample collection used a front and back-up filter sampling approach, in which back-up filters were collected in order to examine if aerosol breakthrough from the front filter occurred or whether evaporation of semivolatiles from the front filter occurred during sampling. In all experiments, no SOA constituents were detected on the back-up filters. Filter sampling was initiated when the aerosol volume reached its maximum (constant) value, as determined by the DMA. Lights were turned off after the filter sampling period was complete. Depending on the total volume concentration of aerosol in the chamber, the duration of filter sampling was 1.8–2.1 h, which resulted in 2.0–2.9 m<sup>3</sup> of total chamber air sampled. Teflon filter extraction protocols in high-purity methanol (LC-MS CHROMASOLV-Grade, Sigma-Aldrich) have been described previously (Surratt et al., 2008). The resultant filter extracts were then analyzed by a Waters ACQUITY ultra performance liquid chromatography (UPLC) system, coupled with a Waters LCT Premier TOF mass spectrometer equipped with an ESI source, allowing for accurate mass measurements by UPLC/ESI-TOFMS to be obtained for each ion (Surratt et al., 2008).

Selected filter extracts from experiments were also analyzed by a Thermo Finnigan Surveyor high performance liquid chromatography (HPLC) system (pump and autosampler) coupled to a Thermo Finnigan LCQ ion trap mass spectrometer (ITMS) equipped with an ESI source, allowing for tandem MS measurements to be obtained. Data were acquired and processed using Xcalibur version 1.3 software. A Waters Atlantis T3 column (3  $\mu$ m particle size; 2.1  $\times$  150 mm) was employed, which is similar to the Water ACQUITY UPLC HSS column used for the UPLC/ESI-TOFMS analysis. The mobile phases consisted of 0.1 % acetic acid in water (A) and 0.1 % acetic acid in methanol (B). The applied 45 min gradient elution program was as follows: the concentration of eluent B was kept at 3 % for 4 min, then increased to 100 % in 21 min, holding at 100 % for 10 min, then decreased to 3 % in 5 min, and kept at 3 %

for 5 min. The injection volume and flow rate were 10  $\mu\text{L}$  and 0.2  $\text{mL min}^{-1}$ , respectively. The ion trap mass analyzer was operated under the following conditions: sheath gas flow ( $\text{N}_2$ ), 65 arbitrary units; auxiliary gas flow ( $\text{N}_2$ ), 3 arbitrary units; source voltage,  $-4.5 \text{ kV}$ ; capillary voltage,  $-14.5 \text{ V}$ ; tube lens offset, 7 V; capillary temperature, 200  $^\circ\text{C}$ ; and maximum ion injection time, 200 ms. Two scan events were used during each chromatographic run; scan event 1 was the full scan mode in which data were collected from  $m/z$  120 to 600 in the negative ionization mode and scan event 2 was the MS2 mode in which product ions were generated from significant base peak ions observed in scan event 1. For MS2 experiments, an isolation width of 2.5  $m/z$  units and a normalized collision energy level of 35 % were applied. The  $[\text{M-H}]^-$  ion signal optimization was carried out by introducing a 1  $\text{mg mL}^{-1}$  malic acid standard solution. Due to the on-axis ESI source that is characteristic of the LCQ ITMS instrument, a solvent delay time of 3.5 min (which diverted the column effluent from the ESI source to waste) was employed to prevent clogging by nonvolatile salts at the entrance of the capillary.

### 2.3.2 High-Resolution Time-of-Flight Aerosol Mass Spectrometry (HR-ToF-AMS)

Real-time aerosol mass spectra were obtained using an Aerodyne HR-ToF-AMS (DeCarlo et al., 2006). The HR-ToF-AMS was operated in a lower resolution, higher sensitivity “V-mode”, and a high-resolution “W” mode, switching between modes once every minute. The “V-mode” data were analyzed to extract sulfate, ammonium, and organic spectra according to procedures in Allan et al. (2004). Calculation of the SOA densities were achieved by comparing the particle mass distributions obtained using the particle ToF mode and the volume distributions obtained by the DMA (Bahreini et al., 2005) in nucleation (seed-free) experiments. O : C, N : C, and H : C ratios were determined from “W” mode data using the APES toolbox and applying the procedures outlined in Aiken et al. (2007) and Aiken et al. (2008). The particle-phase signal of  $\text{CO}^+$  and the organic contribution to  $\text{H}_x\text{O}^+$  ions were estimated as described in Aiken et al. (2008).

Particle-into-Liquid Sampler/Ion Chromatography (PILS/IC) was also employed as described in Kautzman et al. (2010) for the guaiacol experiments, though we find that  $< \text{C}_6$  diacids do not constitute an important fraction of the SOA formed.

## 3 SOA yields

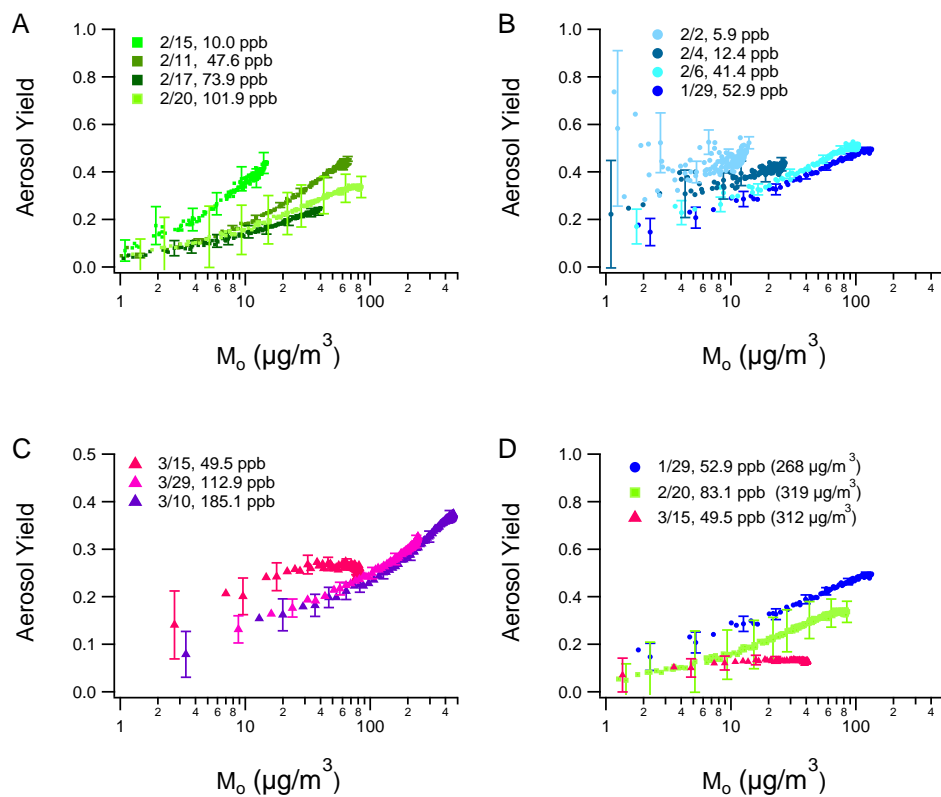
The formation of secondary organic aerosol (SOA) results from the gas-particle partitioning of low-vapor-pressure products formed in the oxidation of volatile organic compounds (VOCs). The SOA yield is defined as the ratio of

mass of organic aerosol formed,  $\Delta M_o$ , to the mass of the parent organic species consumed,  $\Delta \text{Org}$ ,  $Y = \Delta M_o / \Delta \text{Org}$ . The SOA yields for all experiments are summarized in Table 2. To calculate the mass concentration of the SOA, the SOA volumes established by DMA measurements are wall-loss corrected following procedures in Keywood et al. (2004); Ng et al. (2007) and then multiplied by the SOA density, as determined by the AMS in seed-free (nucleation) experiments. A density of 1.65  $\mu\text{g cm}^{-3}$  was determined for phenol. Guaiacol and syringol densities were determined to be 1.45 and 1.49  $\mu\text{g cm}^{-3}$ , respectively. The O : C was calculated at time of maximum SOA growth for these systems and were reported in Chhabra et al. (2011). O : C values of  $0.88 \pm 0.27$ ,  $0.89 \pm 0.28$ , and  $0.97 \pm 0.30$  were calculated for the phenol, guaiacol, and syringol systems, respectively.

Aerosol yields as a function of organic mass for the studied systems are shown in Fig. 1. Phenol traces are denoted by square markers and are of a green shade. Guaiacol traces are denoted by circles and are of a blue shade. Triangle markers indicate the syringol experiments and are colored in various shades of purple. The error bars represent the propagated uncertainty from all parameters varying with time for each experiment. Generally, there is larger error associated with yield measurements at smaller  $M_o$ . Within a system, the yield curves generally converge within uncertainty for  $M_o > 20 \mu\text{g cm}^{-3}$  and there is greater deviation for experiments with the least amount of hydrocarbon reacted within a set of loadings tested for. This may indicate the existence of chemical differences as well as equilibrium and kinetic regimes of aerosol growth that are affected by initial hydrocarbon loading.

About 80 % of the initial phenol was reacted in these experiments. Final SOA yields for phenol range 24–44 % under the experimental conditions (Fig. 1a). For phenol, the hydrocarbon measurement obtained from all gas-phase instruments displayed interferences under low- $\text{NO}_x$  conditions, leading to a wider spread and greater uncertainty in the yield curve parameters. Still, the measured yields under low- $\text{NO}_x$  conditions overlap with the range of 38–45 % reported by Nakao et al. (2011). Greater than 90 % of the initial guaiacol was consumed over the course of all experiments except for the higher guaiacol loading 1/29 experiment as noted in Table 2. SOA yields from guaiacol photooxidation, based on the final aerosol volume achieved, range 44–50 % (Fig. 1b). Within six hours of irradiation, the syringol levels are lower than the limit of detection and SOA yields of 25–37 % result (Fig. 1c). This range overlaps yields of 10–36 % reported for syringol photooxidation under high- $\text{NO}_x$  (approaching 10 ppm NO) conducted by Lauraguais et al. (2012) using  $\text{CH}_3\text{ONO}$  as the OH source.

Figure 1d compares phenol, guaiacol, and syringol experiments in which approximately 300  $\mu\text{g cm}^{-3}$  of hydrocarbon precursor reacted. Phenol and guaiacol follow a similar trend and syringol aerosol is the only system that plateaus with increasing organic aerosol mass, resulting in aerosol



**Fig. 1.** Aerosol yields as a function of organic aerosol mass for (A) phenol (squares), (B) guaiacol (circles), and (C) syringol (triangles). (D) Comparison across systems for similar amounts of reacted hydrocarbon by mass concentration.

yields that are lower than those of phenol and guaiacol at  $M_o > 20 \mu\text{g cm}^{-3}$ .

One might expect that the relatively higher boiling point and lower vapor pressure of syringol might suggest enhanced ability of it or its photooxidation products to partition to the particle-phase, as vapor pressure is a key component in calculating the partitioning coefficient (Pankow, 1994). Syringol vapor-phase wall loss in the reactors is also believed to be minimal. Lauraguais et al. (2012) found no vapor-phase wall loss of syringol and Coeur-Tourneur et al. (2010) found negligible losses for guaiacol and related methoxy and alkyl substituted benzenes. Assuming a first-order wall loss rate,  $k_w \sim 1 \times 10^{-5} \text{ s}^{-1}$ , a typical rate for previously tested compounds in our chamber (Loza et al., 2010), this would lead to roughly a quarter loss of syringol to the walls over the course of the experiment. Correcting for this by lowering  $\Delta\text{Org}$  would lead to syringol SOA yields that are similar to those of guaiacol, but not significantly higher. We further explore the chemical basis for the differences in the yields after examining the gas-phase chemistry.

Most methoxyphenol experiments ended shortly after the hydrocarbon decay was complete. Though, even 4 h after 112.9 ppb syringol was completely reacted in the 3/29 experiment, there was no additional SOA growth. The type of growth observed here is typified by a mechanism involving

oxidation to form SOA either from first-generation products or sufficiently rapid low-volatility product formation over the course of the experiment from further generation reactions (Chan et al., 2007; Ng et al., 2006). We believe it to be the latter explanation in the case of methoxyphenol systems. The high-level of oxidation determined from the measured O : C ratios of the SOA are indicative of multi-generation products, which is consistent with the identified gas and aerosol products discussed below.

#### 4 SOA formation chemistry

We present in this section an analysis of the gas-phase mechanistic chemistry involved in SOA formation based on the CIMS traces for each system along with SOA mass growth over time. All ions unless otherwise noted were monitored during negative mode operation. CIMS signals are plotted in arbitrary units (a.u.). Discussion is focused on the guaiacol system since the off-line aerosol filter analyses complement the gas-phase data. Chemical parallels are drawn between the three systems.

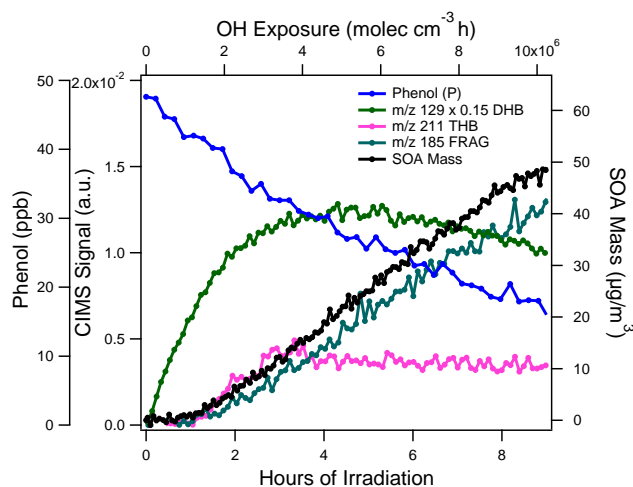
#### 4.1 Gas-phase oxidation

Substituent groups such as  $-\text{OH}$  and  $-\text{OCH}_3$  activate more strongly the aromatic ring towards electrophilic addition of OH compared to the  $-\text{CH}_3$  group. Of these groups, activation potential of the ring increases in order of  $-\text{CH}_3 < -\text{OCH}_3 < -\text{OH}$ . Thus, as predicted, methoxyphenol compounds react faster with OH than their methyl or unsubstituted equivalents (Coeur-Tourneur et al., 2010). Methoxy and hydroxy substituent groups tend to make OH-attack favorable at positions *ortho* and *para* to the substituents, so structures for the measured ions are proposed from chemical mechanisms assuming one of these positions of initial OH-attack. H-atom abstraction from the methoxy group is expected to be small, as this path was determined to contribute  $< 4\%$  in the case of methoxybenzene (Coeur-Tourneur et al., 2010).

A generation is defined as the OH-initiated reaction of a stable (non-radical) species, and the OH exposure is calculated as the product of the OH concentration (inferred from the parent hydrocarbon decay) and the hours of irradiation. Each  $m/z$  is the sum of the signal from all isomeric structures detected by the CIMS at that ion, but are not shown explicitly in the abbreviated mechanisms and tables presented. However, from the chemical ionization method employed, we expect certain common chemical features of the proposed structures. For example, many of the transfer  $[\text{M}+19]^-$  products are likely acidic (Huey et al., 1996), containing carboxylic acid groups and/or sufficient acidic hydroxyl groups (Tables 3, 4, 6).

#### 4.2 Phenol chemistry

Initial steps of the phenol gas-phase mechanism have been elucidated elsewhere (Olariu et al., 2002; Berndt and Boge, 2003). Phenol + OH proceeds primarily with *ortho*-OH addition to the ring,  $\text{O}_2$  addition, followed by elimination of  $\text{HO}_2$  to form 1,2-dihydroxybenzene (catechol). In the low- $\text{NO}_x$  phenol experiments, CIMS ions are tracked that show successive OH adduct product formation up to three generations. Quinone products are also likely, though these are not detected in negative mode operation of the current CIMS technique unless they are additionally functionalized. A list of ions monitored in this system is presented in Table 3. Representative data for the phenol low- $\text{NO}_x$  system are shown in Fig. 2. Phenol,  $\text{C}_6\text{H}_6\text{O}$ , is monitored at  $m/z$  179, a cluster product. We do not include  $m/z$  113, the fluoride transfer product, because this signal includes interference from another compound (likely a small acid) that grows in over time. Dihydroxybenzene,  $\text{C}_6\text{H}_6\text{O}_2$ , is observed primarily as the fluoride transfer product  $m/z$  129 (DHB) with less signal observed at the  $m/z$  195 cluster product. A trace at  $m/z$  211 indicates a third OH addition to the ring to form trihydroxybenzene (THB),  $\text{C}_6\text{H}_6\text{O}_3$ . These time profiles are shown along with SOA growth and phenol decay for the first nine hours





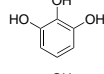
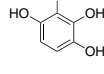
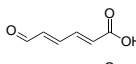
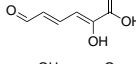
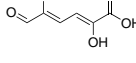


**Fig. 2.** Phenol low- $\text{NO}_x$  gas-phase and particle-phase development. The signal at  $m/z$  129 was scaled by a factor of 0.15.

of irradiation (Fig. 2). SOA growth occurs coincident with the growth of  $m/z$  211 (THB), indicating that the transition to lower volatility products to form SOA likely occurs along the 3rd addition of OH just after  $\sim 1 \times 10^6$  molec  $\text{cm}^{-3}$  h of OH exposure. This is the approximate equivalent of one hour of photooxidation in the atmosphere assuming an ambient OH concentration of  $\sim 1 \times 10^6$  molec  $\text{cm}^{-3}$ . A signal at  $m/z$  227 was observed to have a similar profile as  $m/z$  211, and may be attributed to tetra hydroxybenzene (4HB, Table 3) and an epoxide isomer ( $\text{C}_6\text{H}_6\text{O}_4$ ). A plateau in the profile is consistent with 4HB also contributing to the particle phase. Individual contributions from 4HB and the epoxide could not be determined from the signal at  $m/z$  227, which was approximately one-fourth of the signal at  $m/z$  211 (not shown for clarity). A minor amount of signal at  $m/z$  229 is attributed to the hydroperoxide  $\text{C}_6\text{H}_8\text{O}_4$ .

Figure 2 also shows that the SOA is in equilibrium with  $m/z$  185, which we propose to be the ring fragment  $\text{C}_4\text{H}_4\text{O}_3$ , hydroxy butenedial, that forms from the decomposition of the bicyclic radical from phenol + OH. Another dicarbonyl,  $\text{C}_4\text{H}_4\text{O}_4$ , is monitored at  $m/z$  135, which could be the analogous ring fragment from dihydroxybenzene + OH. These products approach the high O:C ratios characteristic of the SOA in this system. Comparisons of the CIMS gas-phase traces and the off-line filter analyses from similar experiments performed by Nakao et al. (2011) provide additional insights. Nakao et al. (2011) utilized electrospray ionization and atmospheric pressure chemical ionization mass spectrometry (ESI/APCI-TOFMS) as well as PILS-ESI-TOFMS for analysis, though Nakao et al. (2011) noted that the PILS-ESI-TOFMS spectra may include water soluble gas-phase species because a denuder was not used. While Nakao et al. (2011) observed an exact mass match that corresponds with the bicyclic hydroperoxide from phenol ( $\text{C}_6\text{H}_8\text{O}_6$ ) and another compound with a suggested formula  $\text{C}_6\text{H}_8\text{O}_7$  using



**Table 3.** Proposed structures for CIMS ions in the phenol low-NO<sub>x</sub> system. C and T indicate the cluster and transfer product, respectively.

Observed $m/z$	Product	Chemical Formula	Proposed Structure (one isomer shown)	Chemical Pathway
179	C	C <sub>6</sub> H <sub>6</sub> O		phenol
129	T <sup>a</sup>	C <sub>6</sub> H <sub>6</sub> O <sub>2</sub>		DHB = phenol + OH
211	C	C <sub>6</sub> H <sub>6</sub> O <sub>3</sub>		THB = DHB + OH
227	C <sup>b</sup>	C <sub>6</sub> H <sub>6</sub> O <sub>4</sub>		4HB = THB + OH
145	T	C <sub>6</sub> H <sub>6</sub> O <sub>3</sub>		ring opening acid
161	T	C <sub>6</sub> H <sub>6</sub> O <sub>4</sub>		ring opening acid
177	T	C <sub>6</sub> H <sub>6</sub> O <sub>5</sub>		ring opening acid
185	C <sup>c</sup>	C <sub>4</sub> H <sub>4</sub> O <sub>3</sub>		ring fragment
135	T <sup>d</sup>	C <sub>4</sub> H <sub>4</sub> O <sub>4</sub>		ring fragment

<sup>a</sup>  $m/z$  195 is also present as C<sub>6</sub>H<sub>6</sub>O<sub>2</sub>, but is lower signal than at  $m/z$  129

<sup>b</sup> Signal at  $m/z$  227 likely has an epoxide contribution as well

<sup>c</sup> The transfer product at C<sub>4</sub>H<sub>4</sub>O<sub>3</sub> is at  $m/z$  119, which is not monitored in low-NO<sub>x</sub> conditions because it is the cluster product of H<sub>2</sub>O<sub>2</sub>

<sup>d</sup>  $m/z$  201 is also present as C<sub>4</sub>H<sub>4</sub>O<sub>4</sub>, but is lower signal than at  $m/z$  135

both techniques, we do not observe these products in the gas-phase with the CIMS. It is possible that these compounds are of sufficiently low volatility at this point that they are not measurable in the gas phase. The CIMS does detect ions corresponding to several < C<sub>6</sub> products seen in Nakao et al. (2011), including: C<sub>4</sub>H<sub>4</sub>O<sub>3</sub> at  $m/z$  185 and C<sub>4</sub>H<sub>4</sub>O<sub>4</sub> at  $m/z$  135 and  $m/z$  201.

Nakao et al. (2011) also observed a series of oxygen additions, from C<sub>6</sub>H<sub>6</sub>O<sub>2</sub> to C<sub>6</sub>H<sub>6</sub>O<sub>5</sub>, which are interpreted to possibly be a series of OH adduct products. Another structural possibility is that these are C<sub>6</sub> retaining, but ring-opened carboxylic acids that would correspond to large signals in the CIMS as transfer products:  $m/z$  145,  $m/z$  161, and  $m/z$  177. From the time profiles of these CIMS ions, we predict that they are likely carboxylic acids even though the proposed chemical formulae are isomeric to the aromatic OH adducts. That is, these ion signals grow linearly with time (not shown) during the experiment duration and do not peak or plateau like their isomeric aromatic OH adducts. The peak/plateau of the signal from aromatic OH adducts is consistent with the proposed products' functionalities and potential to partition into the particle-phase, whereas these ring-opened products are of higher vapor pressure and may form throughout

the experiment during continued oxidation of the phenolic products.

The presence of these products may partially explain the absence of a strong signal from a hydroperoxide C<sub>6</sub>H<sub>8</sub>O<sub>4</sub> at  $m/z$  229. The alkylperoxy radical preferentially isomerizes to form the bicyclic peroxy radical and decomposes from a bicyclic alkoxy radical to < C<sub>6</sub> fragments (often dicarbonyl) as mentioned above. However, Birdsall et al. (2010) observed evidence that the alkylperoxy radical may also participate in chemistry that regenerates OH and opens the ring (alternate route to alkoxy radical formation without first forming the bicyclic peroxy radical). This was found to be a more minor pathway also via RO<sub>2</sub> + HO<sub>2</sub> → RO + OH + O<sub>2</sub> for the case of toluene photooxidation. This results in an aromatic alkoxy radical and ring-opening to form methylhexadienedial, a C<sub>7</sub> dialdehyde (preserving the carbon backbone). Though this route was more prominent under low [O<sub>2</sub>] conditions, Birdsall et al. (2010) also observed this product at pressures more similar to the atmosphere where bicyclic peroxy radical formation is expected to dominate and form dicarbonyl ring fragments. Thus, this ring-opening dicarbonyl pathway may also be present in this study, while minor compared to production of dicarbonyl fragments.



**Table 4.** Proposed structures for CIMS ions in the guaiacol low-NO<sub>x</sub> system. C and T indicate the cluster and transfer product, respectively.

Observed $m/z$	Product	Chemical Formula	Proposed Structure (one isomer shown)	Chemical Pathway <sup>a</sup>
209	C <sup>b</sup>	C <sub>7</sub> H <sub>8</sub> O <sub>2</sub>		guaiacol
259	C	C <sub>7</sub> H <sub>10</sub> O <sub>5</sub>		(1) hydroperoxide
225	C	C <sub>7</sub> H <sub>8</sub> O <sub>3</sub>		(2) G + OH = guaiacol + OH
241	C	C <sub>7</sub> H <sub>8</sub> O <sub>4</sub>		(2) G + 2OH = guaiacol + 2OH
191	T <sup>c</sup>	C <sub>7</sub> H <sub>8</sub> O <sub>5</sub>		(2) G + 3OH = guaiacol + 3OH
175	T	C <sub>7</sub> H <sub>8</sub> O <sub>4</sub>		ring opening product
257	C <sup>d</sup>	C <sub>7</sub> H <sub>8</sub> O <sub>5</sub>		(3aii) EPOX and (2) G + 3OH
149	T	C <sub>5</sub> H <sub>6</sub> O <sub>4</sub>		(3aiii) ring fragment
187	C	C <sub>4</sub> H <sub>6</sub> O <sub>3</sub>		(3aiii) ring fragment
199	C	C <sub>5</sub> H <sub>6</sub> O <sub>3</sub>		(3aiii) ring fragment
107	T	C <sub>3</sub> H <sub>4</sub> O <sub>3</sub>		(3aiii) ring fragment
189	C	C <sub>3</sub> H <sub>4</sub> O <sub>4</sub>		(3aiii) ring fragment
243	T	C <sub>6</sub> H <sub>6</sub> O <sub>5</sub>		(3b) C <sub>6</sub> bicyclic ketone
129	T <sup>e</sup>	C <sub>6</sub> H <sub>6</sub> O <sub>2</sub>		(4) Methoxy Loss

<sup>a</sup> Numbers indicate correspondence with pathways outlined in mechanism (Fig. 4); <sup>b</sup>  $m/z$  143 is also present as C<sub>7</sub>H<sub>8</sub>O<sub>2</sub>, but is lower signal than at  $m/z$  209; <sup>c</sup> Also includes signal from C<sub>7</sub>H<sub>8</sub>O<sub>5</sub> ring opening acid; <sup>d</sup> Includes contributions from G + 3OH along pathway 2 in Fig. 4; <sup>e</sup>  $m/z$  195 is also present as C<sub>6</sub>H<sub>6</sub>O<sub>2</sub>, but is lower signal than at  $m/z$  129

Figure 3 outlines a potential mechanism employing this scheme to form these multifunctional C<sub>6</sub> carboxylic acids from phenol, dihydroxybenzene, and trihydroxybenzene. The difference between the products proposed here and those of Birdsall et al. (2010) is that the hydroxy group placement

results in carboxylic acid functionality. We note that depending on the hydroxy group placement on the ring, isomeric hydroxy dicarbonyl products are also likely and can contribute to these ion signals as well. OH regeneration could also come from photolysis of a hydroperoxide producing a favorable

**Table 5.** SOA products observed in UPLC/(-)ESI-TOFMS offline filter analysis for the guaiacol low-NO<sub>x</sub> system.

[M-H] <sup>-</sup>	Meas. Mass	Suggested Chemical Formula (M)	Error* (mDa)
189	189.0378	C <sub>7</sub> H <sub>10</sub> O <sub>6</sub>	-2.1
219	219.014	C <sub>7</sub> H <sub>8</sub> O <sub>8</sub>	-0.1
147	147.0266	C <sub>5</sub> H <sub>8</sub> O <sub>5</sub>	-2.6
157	157.0113	C <sub>6</sub> H <sub>6</sub> O <sub>5</sub>	-2.4
133	133.0133	C <sub>4</sub> H <sub>6</sub> O <sub>5</sub>	-2.4
149	149.0072	C <sub>4</sub> H <sub>6</sub> O <sub>6</sub>	-1.4
205	205.0327	C <sub>7</sub> H <sub>10</sub> O <sub>7</sub>	2.1
145	145.048	C <sub>6</sub> H <sub>10</sub> O <sub>4</sub>	-2.1
163	163.0233	C <sub>5</sub> H <sub>8</sub> O <sub>6</sub>	-1
203	203.0177	C <sub>7</sub> H <sub>8</sub> O <sub>7</sub>	-1.5
171	171.0279	C <sub>7</sub> H <sub>8</sub> O <sub>5</sub>	-1.4
169	169.0124	C <sub>7</sub> H <sub>6</sub> O <sub>5</sub>	-1.3
173	173.0428	C <sub>7</sub> H <sub>10</sub> O <sub>5</sub>	-2.2
115	115.0004	C <sub>4</sub> H <sub>4</sub> O <sub>4</sub>	-2.7
129	129.0175	C <sub>5</sub> H <sub>6</sub> O <sub>4</sub>	-1.3

\* The error in the accurate mass measurements is the difference between the theoretical mass of the suggested molecular formulae and the measured mass of the ion.

alkoxy radical for ring opening, though we would not expect this to be significant on such a quick timescale that is consistent with negligible signal from the hydroperoxide. All CIMS traces for proposed ring opening acids and ring fragments share a similar trend of constant linear growth over time (not shown). Many of these trend with the aerosol mass (e.g. *m/z* 185 mentioned earlier), indicating that aerosol growth continues to proceed with the generation of ring opening and fragmentation processes.

### 4.3 Guaiacol chemistry

Figure 4 presents proposed oxidation pathways for the guaiacol system under low-NO<sub>x</sub> conditions. Solid boxed structures indicate that the expected *m/z* from the chemical ionization reactions in the CIMS for that compound was detected. The *m/z* monitored is indicated, as are proposed C<sub>x</sub>H<sub>y</sub>O<sub>z</sub> formulas and molecular weights. A list of ions monitored in this system is presented in Table 4. Chemical formulas shown in red indicate correspondence with a suggested C<sub>x</sub>H<sub>y</sub>O<sub>z</sub> for accurate mass observations in filter data (Table 5). The particular structures in Fig. 4 indicate potential structures for the accurate mass suggested chemical formulae and the mechanism does not explicitly present all potential isomeric structures. The boxed colors correspond with the colors used for ion time traces as measured by the CIMS and included in Fig. 5.

Guaiacol gas-phase oxidation proceeds with OH addition to the ring, followed by reaction with oxygen to form an organic peroxy radical. Under the experimental conditions, the fate of this RO<sub>2</sub> radical is reaction with HO<sub>2</sub> to form a hy-

droperoxide (Fig. 4, Pathway 1), elimination of HO<sub>2</sub> to retain aromaticity (Fig. 4, Pathway 2), or isomerization to the bicyclic radical (Fig. 4, Pathway 3).

#### 4.3.1 Pathway 1: RO<sub>2</sub> + HO<sub>2</sub>

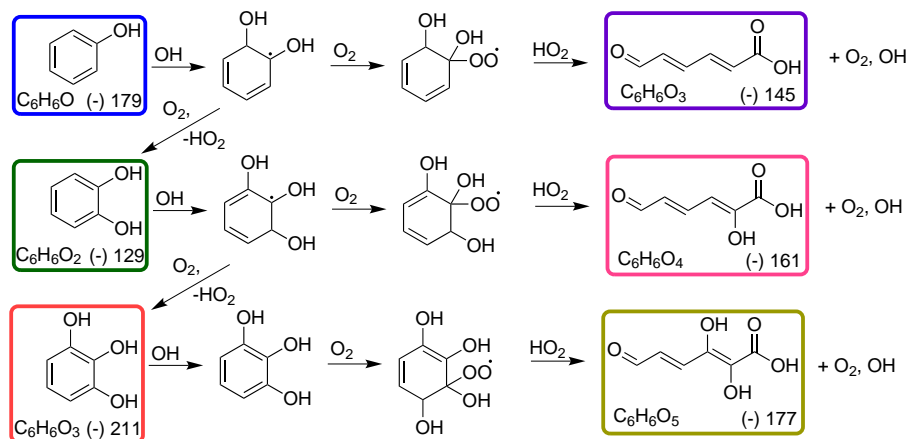
Despite predominant conditions of RO<sub>2</sub> + HO<sub>2</sub> chemistry, we do not observe strong signals in the CIMS measurements for potential hydroperoxides, as is similar to the case of phenol. A small trace at *m/z* 259 suggests that there is slight production of the hydroperoxide C<sub>7</sub>H<sub>10</sub>O<sub>5</sub>. The off-line filter analysis does reveal a small contribution from a corresponding accurate mass measurement (Table 5). More significant signals in the CIMS are presented in Fig. 5, demonstrating that the preferred routes are to preserve aromaticity or involve isomerization to the bicyclic radical. This observation is consistent with previous studies on aromatic systems (Calvert et al., 2002; Bloss et al., 2005; Johnson et al., 2005; Birdsall et al., 2010).

#### 4.3.2 Pathway 2: continuous OH addition to ring

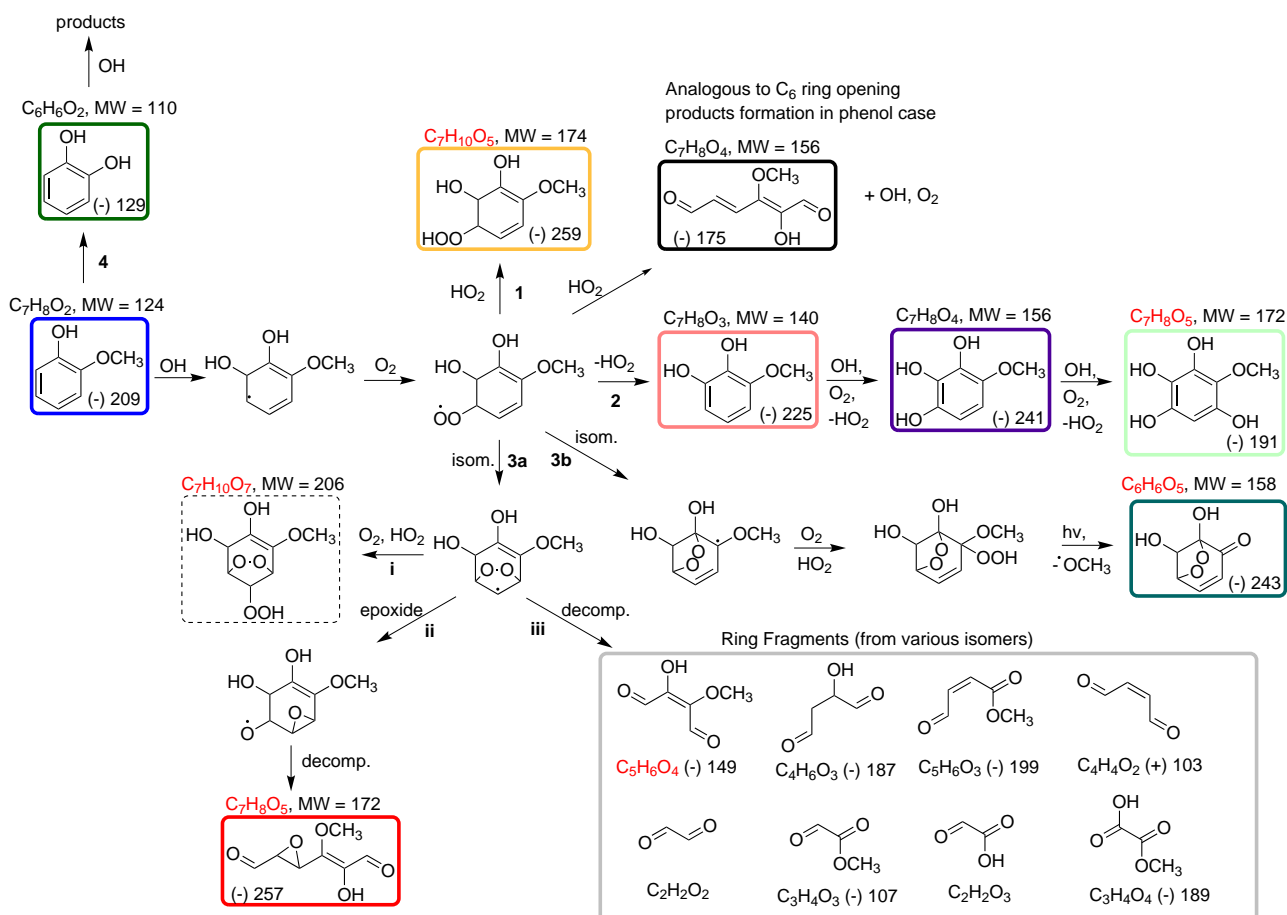
The general chemical development of the consecutive OH addition pathways in the guaiacol low-NO<sub>x</sub> system are tracked in Fig. 5. Figure 4, Pathway 2 is marked by *m/z* 225 on the CIMS, a cluster product for the guaiacol OH-adduct product, C<sub>7</sub>H<sub>8</sub>O<sub>3</sub> and denoted as G + OH. Successive OH addition products are monitored at 16 amu increments at *m/z* 241 and *m/z* 257, for C<sub>7</sub>H<sub>8</sub>O<sub>4</sub> denoted as G + 2OH and C<sub>7</sub>H<sub>8</sub>O<sub>5</sub> denoted as G + 3OH, respectively. The third OH adduct can be tracked at both the transfer and the cluster products, *m/z* 191 and *m/z* 257, respectively. However, *m/z* 257 includes the combined signal of rapid epoxide formation and the later formed third generation OH adduct. A signal at *m/z* 191 is predominantly the third generation OH adduct during the first 8 h until another product (likely an acid) grows in at this *m/z*. The first OH-adduct product (*m/z* 225) is short-lived due to further activation of the ring by the additional hydroxy group. Formation of C<sub>7</sub>H<sub>8</sub>O<sub>4</sub> monitored at *m/z* 241 follows promptly. The signal at *m/z* 191 (not shown) was observed to follow after *m/z* 241, indicating that third OH addition is achieved. C<sub>7</sub>H<sub>8</sub>O<sub>5</sub> is also a proposed chemical formula for a product observed in the filter analysis (Table 5), suggesting that the third OH adduct is of sufficiently low volatility to partition to the particle phase.

#### 4.3.3 Pathway 3: isomerization to bicyclic radical

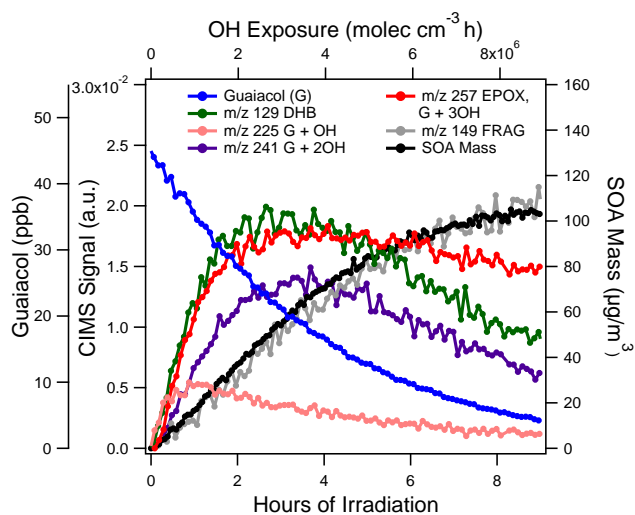
The bicyclic peroxy radical can react with HO<sub>2</sub> to form the bicyclic hydroperoxide C<sub>7</sub>H<sub>10</sub>O<sub>7</sub> (Fig. 4, Pathway 3ai), however, no evidence exists for this product in the CIMS. This is analogous with the bicyclic hydroperoxide from the phenolic case. However, if this product is formed, it already has a low vapor pressure and an O : C of 1, and would likely partition into the aerosol phase. Filter data also indicate the presence of C<sub>7</sub>H<sub>10</sub>O<sub>7</sub> in the particle-phase (Table 5). The majority of



**Fig. 3.** Proposed pathway for gas-phase production of  $C_6$  carboxylic acids in the phenol low- $NO_x$  system. Boxed structures indicate that the expected  $m/z$  from the chemical ionization reactions in the CIMS for that compound was detected.



**Fig. 4.** Proposed mechanism for guaiacol low- $NO_x$  oxidation. Boxed structures indicate that the expected  $m/z$  from the chemical ionization reactions in the CIMS for that compound was detected. Chemical formulae in red correspond with those proposed for the filter accurate mass aerosol measurements.

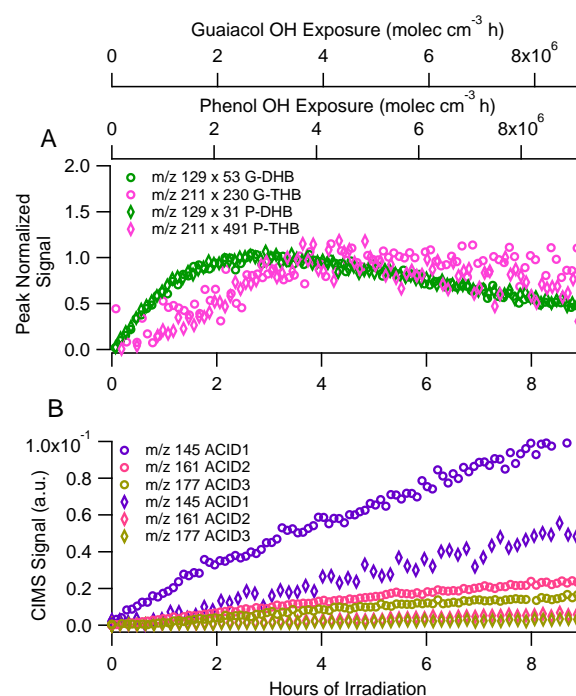


**Fig. 5.** Selected traces in guaiacol low- $\text{NO}_x$  gas-phase and particle-phase development.

the filter data guaiacol products exhibit O : C ratio  $> 0.7$ , and the average bulk O : C ratio as measured by the AMS is 0.89, so it is likely these hydroperoxide species may immediately partition and become further processed in the particle phase. Ofner et al. (2011) found O : C ratios of 0.3–1, which were attributed to products that are likely highly oxygenated benzenes and conjugated olefins.

Alternatively, the bicyclic radical can rearrange, breaking the oxygen bridge to form an epoxide and open the ring (Fig. 4, Pathway 3aii). There is evidence of a significant epoxide (EPOX) route that grows in early with the dihydroxybenzene route (Fig. 5). The epoxide is monitored at  $m/z$  257, the cluster product of  $\text{C}_7\text{H}_8\text{O}_5$  (Fig. 5), though there is some later contribution from the third generation OH adduct, as mentioned earlier. The epoxide will continue to react with OH to generate more functionalized molecules that will likely contribute to SOA formation.

The bicyclic radical can also decompose to several fragments as shown in the box in Fig. 4, Pathway 3aiii. A major fragment is observed at  $m/z$  149, which could be the methoxy analog to the phenolic fragment, hydroxybutenedial ( $\text{C}_4\text{H}_4\text{O}_3$ ) at  $m/z$  185. We propose that  $m/z$  149 is a methoxy-hydroxy-butenedial fragment,  $\text{C}_5\text{H}_6\text{O}_4$ .  $\text{C}_5\text{H}_6\text{O}_4$  is also an accurate mass assignment for a product observed in the filter data (Table 5). The  $m/z$  149 trace has a similar profile to the SOA mass curve, indicating that it is in equilibrium with the particle phase (Fig. 5). Other possible fragments are observed at ions indicated in the mechanism. The  $< \text{C}_4$  fragments approach O : C values  $> 1$ , and the  $\text{C}_4$  and  $\text{C}_5$  fragments would likely undergo further oxidation in the gas phase and then possibly contribute to the particle phase. The fate of butenedial ( $\text{C}_4\text{H}_4\text{O}_2$ ) has been described elsewhere (Calvert et al., 2002; Bloss et al., 2005).



**Fig. 6.** Common chemical routes in the phenol and guaiacol low- $\text{NO}_x$  systems. (A) Dihydroxybenzene routes in phenol (diamonds) and guaiacol (open circles). Each CIMS trace is normalized by the peak value. (B) Common acids in phenol (diamonds) and guaiacol (open circles).

#### 4.3.4 Pathway 4: methoxy loss route

The early growth of a strong signal at  $m/z$  129 (Fig. 5) and a matching time profile at  $m/z$  195 upon irradiation suggests that dihydroxybenzene ( $\text{C}_6\text{H}_6\text{O}_2$ ) is somehow formed from guaiacol ( $\text{C}_7\text{H}_8\text{O}_3$ ) + OH in the gas phase. This is further supported by matching time profiles from the guaiacol and phenol systems for dihydroxybenzene (DHB) and its oxidation to trihydroxybenzene (THB) in Fig. 6a, as well as growth of the proposed  $\text{C}_6$  ring opening acids (Fig. 6b). In the guaiacol system, the raw signal at  $m/z$  211 includes contributions from the  $^{18}\text{O}$  natural isotope peak of the guaiacol cluster product and trihydroxybenzene. The signal at  $m/z$  211 has been corrected to reflect only the contribution from trihydroxybenzene (G-THB). There is negligible signal at  $m/z$  227 in the guaiacol system, indicating that the dihydroxybenzene that is formed proceeds only to one more step of oxidation before the carbon along this route is incorporated into pathways besides continuous OH adduct formation. Rather, the carbon likely ends up in production of ring-opened acids as shown in Fig. 6b.

A trace at  $m/z$  243 is present in the guaiacol system, though because there is negligible signal at  $m/z$  227, it is unlikely that this ion is part of the OH adduct series from dihydroxybenzene as  $\text{C}_6\text{H}_6\text{O}_5$ . Still, an exact mass match for  $\text{C}_6\text{H}_6\text{O}_5$  is the fourth highest product (in terms of area counts) of

constituents identified in the filter data, and  $m/z$  243 follows the SOA trace indicating that it may be an important contributor to the particle phase. In the event that the radical placement occurs on the carbon containing the methoxy group during formation of the bicyclic radical, carbon loss may occur by removal of a methoxy group. The ether alkyl radical will react with  $O_2$  to generate an ether alkyl peroxy radical. Collins et al. (2005) proposed that photooxidation of diisopropyl ether in the absence of  $NO_x$  (with  $RO_2 + HO_2$  reactions dominating) likely resulted in the formation of hydroperoxide species that could then undergo decomposition, photolysis, or reaction with OH. Photolysis of the hydroperoxy group in the current methoxyphenol case would result in an alkoxy radical and elimination of the methoxy group to form a ketone. This leads to formation of a  $C_6$  bicyclic ketone product as outlined in Pathway 3b of Fig. 4. The mechanism and structure for this particle-phase product remain tentative, as the proposed structure may be unstable and participate in additional condensation reactions.

#### 4.3.5 Guaiacol aerosol-phase products

The off-line filter analysis suggests that there are additional chemical pathways for this system than those inferred from the gas-phase CIMS measurements. The top fifteen ions (by highest peak area) are presented in Table 5 along with measured masses and suggested corresponding chemical formulae. Of the suggested chemical formulae, the O : C ranges from 0.67 to 1.5. The  $< C_7$  components are likely ring fragments that have undergone further oxidation to achieve such high O : C compared to those shown in Fig. 4, Pathway 3a.iii. Of the  $C_7$  retaining products listed in Table 5, all suggested chemical formulae except  $C_7H_5O_5$  can be explained by first-generation products along pathways 1, 2, 3ai, and 3a.ii (Fig. 4) and the analogous pathways in higher generation chemistry from the OH adducts. Similar structures to those proposed here have been observed previously in EI-MS analyses (Justesen, 2001). However, since many corresponding signals are not observed in the CIMS, it is also possible that a number of these highly oxygenated ( $> O_5$ ) species may be formed in the particle phase, as evidenced in Ofner et al. (2011) for aerosol from guaiacol and catechol.

Previous study by Ofner et al. (2011) reports that the most important functional groups of generated SOA from guaiacol and catechol include carboxylic acids, carboxylic anhydrides and lactones or esters. While carboxylic acids can form as ring fragmentation products depending on the particular conformation of the bicyclic radical, the observation of carboxylic anhydrides and lactones suggests potential for additional intra- and inter-molecular reactions that may occur in the particle-phase involving the carbonyl and methoxy groups of SOA products. Ofner et al. (2011) also observe the degradation of the methoxy group during SOA formation from guaiacol, and finds this degradation to be more pronounced under wet conditions. This supports additional

chemistry (likely in the particle phase) that goes beyond the proposed gas-phase products in Fig. 4.

#### 4.4 Syringol chemistry

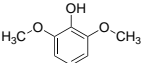
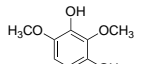
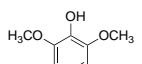
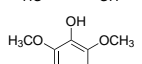
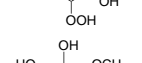
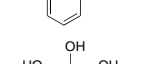
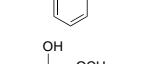
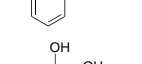
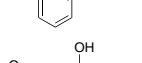
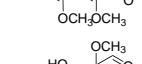
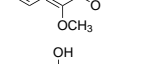
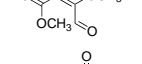
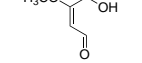
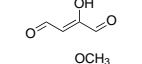
The chemistry of syringol under low- $NO_x$  conditions leads to rapid formation of a rich diversity of gas-phase products and SOA growth. Whereas at least two generations of OH oxidation of phenol are required to generate SOA, guaiacol and syringol produce SOA in the first generation. This is clear when comparing the SOA mass curves between systems in Figs. 2, 5, and 7. As expected, syringol is more reactive than guaiacol because of its additional methoxy group. Syringol is detected primarily at the transfer  $[M+19]^-$ ,  $m/z$  173, with some signal at the cluster product  $[M+85]^-$ ,  $m/z$  239. This is the opposite of guaiacol, the signal of which is found predominantly at the cluster product. This suggests that the extra methoxy group on syringol compared to guaiacol leads to an enhancement towards greater acidity to be seen predominantly at the transfer product ( $m/z$  173). This shift is observed for many of the ions in the syringol case that are analogous to the guaiacol system. CIMS ions monitored in this system are presented in Table 6.

Syringol oxidation chemistry seems to undergo similar routes as guaiacol with analogous ions monitored by the CIMS for the following pathways: (1) hydroperoxide formation, (2) OH adduct formation up to two to three generations, and (3) epoxide and  $< C_8$  ring fragment formation from decomposition of the bicyclic radical. These pathways are highlighted in Fig. 7a by selected ions at 189 and 271 for OH adducts  $C_8H_{10}O_4$  (S + OH) and  $C_8H_{10}O_5$  (S + 2OH), respectively,  $m/z$  287 for the  $C_8H_{10}O_6$  syringol epoxide (SE-POX), and a  $C_6H_8O_5$  ring fragment (FRAG) at  $m/z$  245. A minor contribution from the  $C_8H_{12}O_6$  syringol hydroperoxide at  $m/z$  289 was observed in the CIMS spectra and is not shown.

These routes seem to be minor, however, in comparison to those that favor immediate scission of the  $C_8$  backbone (Fig. 7). Many of the higher signals in the CIMS are at  $m/z$  values that can only be reasonably explained by  $< C_8$  molecular formulas. For example, major traces are from ions at  $m/z$  275 (Fig. 7b), likely  $C_7H_{10}O_6$ . A signal was also observed in the CIMS spectra at  $m/z$  291, possibly  $C_7H_{10}O_7$ , and is not shown. There is also evidence of methoxy and hydroxy exchanges and eliminations, providing even more diversity of masses in the CIMS spectra from syringol photooxidation. More detailed analyses on chemical structure would be required to confirm precise mechanisms.

For syringol, the immediate loss of carbon through novel chemical pathways (methoxy group elimination or exchange with hydroxy) is more evident in multiple routes than was observed in the case of guaiacol. Figure 8 proposes several possibilities of syringol conversion to compounds with fewer methoxy groups. Each of these pathways has representative

**Table 6.** Proposed structures for CIMS ions in the syringol low-NO<sub>x</sub> system. C and T indicate the cluster and transfer product, respectively.

Observed $m/z$	Product	Chemical Formula	Proposed Structure (one isomer shown)	Chemical Pathway <sup>a</sup>
173	T <sup>b</sup>	C <sub>8</sub> H <sub>10</sub> O <sub>3</sub>		syringol
189	T	C <sub>8</sub> H <sub>10</sub> O <sub>4</sub>		S + OH = syringol + OH
271	C	C <sub>8</sub> H <sub>10</sub> O <sub>5</sub>		S + 2OH = syringol + 2OH
289	C	C <sub>8</sub> H <sub>12</sub> O <sub>6</sub>		hydroperoxide
225	C	C <sub>7</sub> H <sub>8</sub> O <sub>3</sub>		(a) G + OH = guaiacol + OH
211	C	C <sub>6</sub> H <sub>6</sub> O <sub>3</sub>		(b) THB = DHB + OH
209	C	C <sub>7</sub> H <sub>8</sub> O <sub>2</sub>		(c) guaiacol
129	T <sup>c</sup>	C <sub>6</sub> H <sub>6</sub> O <sub>2</sub>		(d) DHB = phenol + OH
205	T <sup>d</sup>	C <sub>8</sub> H <sub>10</sub> O <sub>5</sub>		ring opening product
287	C	C <sub>8</sub> H <sub>10</sub> O <sub>6</sub>		SEPOX
245	C	C <sub>6</sub> H <sub>8</sub> O <sub>5</sub>		ring fragment
149	T	C <sub>5</sub> H <sub>6</sub> O <sub>4</sub>		ring fragment
185	C	C <sub>4</sub> H <sub>4</sub> O <sub>3</sub>		ring fragment
203	C	C <sub>4</sub> H <sub>6</sub> O <sub>4</sub>		ring fragment
275	C	C <sub>7</sub> H <sub>10</sub> O <sub>6</sub>	none proposed	C <sub>8</sub> scission
291	C	C <sub>7</sub> H <sub>10</sub> O <sub>7</sub>	none proposed	C <sub>8</sub> scission

<sup>a</sup> Letters indicate correspondence with pathways outlined in mechanism (Fig. 8).

<sup>b</sup>  $m/z$  239 is also present as C<sub>8</sub>H<sub>10</sub>O<sub>3</sub>, but is lower signal than at  $m/z$  173.

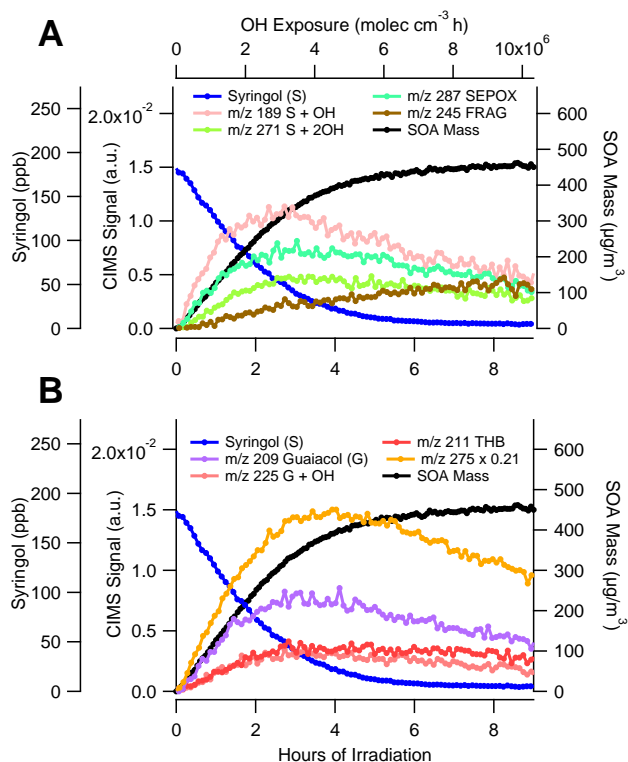
<sup>c</sup>  $m/z$  195 is also present as C<sub>6</sub>H<sub>6</sub>O<sub>2</sub>, but is lower signal than at  $m/z$  129.

<sup>d</sup> Includes contributions from S + 2OH as well as ring opening product.

ions that also appear in the guaiacol case, though they are not necessarily the same mechanistically.

Analogous to the formation of dihydroxybenzene in the case of guaiacol + OH, the early growth of  $m/z$  225 (presumably C<sub>7</sub>H<sub>8</sub>O<sub>3</sub>) suggests that hydroxy guaiacol (G + OH) may form from syringol + OH (Figs. 7b, 8, Pathway a). Another possibility is the complete exchange of both methoxy groups for hydroxy groups, resulting in trihydroxybenzene (THB)

monitored at  $m/z$  211 (Figs. 7b, 8, Pathway b). Formation of guaiacol is also evident by matching time profiles growing in at  $m/z$  143 and 209 (Figs. 7b, 8, Pathway c). Growth of dihydroxybenzene (DHB) at  $m/z$  129 and  $m/z$  195 is observed later in the experiment, likely from conversion of guaiacol (Fig. 8, Pathway d). Observation of signal at  $m/z$  257 (the guaiacol epoxide and third OH adduct) supports that there is a guaiacol channel in syringol photooxidation, and is not



**Fig. 7.** Selected traces in syringol low-NO<sub>x</sub> gas-phase and particle-phase development. **(A)** OH addition pathways. **(B)** Traces indicating methoxy loss pathways.

shown. Though a guaiacol channel exists, the coincident formation of *m/z* 225 (G + OH) and *m/z* 211 (THB) with that of guaiacol at *m/z* 209 is consistent with their direct formation from syringol + OH.

#### 4.5 Novel chemical pathways for loss of carbon in methoxyphenol systems

We observe evidence that guaiacol and syringol OH-initiated photooxidation leads to loss of carbon, likely due to unique chemistry involving the methoxy group. While gas-phase aromatic chemistry is not completely understood (Calvert et al., 2002), we propose possible explanations for these carbon loss pathways. Carbon loss from aromatic systems has been observed previously in the gas-phase, and is thought to be a result of *ipso* addition of OH at a substituent containing ring carbon and subsequent elimination of the substituent leaving a phenolic group. This is the case for alkylbenzenes, in which Noda et al. (2009) observed phenol formation from OH-initiated photooxidation of toluene and cresol formation from xylenes. Prior, Koch et al. (2007) provided evidence that the *ipso* addition of OH is relevant for the case of hexamethylbenzene. The extent by which these processes occur in the gas-phase is still unclear, as Aschmann et al. (2010) does not observe the dealkylation as in Noda et al. (2009). Addition at the *ipso* position is generally considered to be a minor

route ( $\leq 20\%$  for the alkylbenzenes studied in Noda et al., 2009), though it could be that oxy substituents may influence the favorability of this route. Raff and Hites (2006) found that brominated diphenylethers also undergo ether cleavage, proposed as a result of *ipso* addition of OH. Thus, *ipso* addition of OH in the gas-phase may be more relevant for more highly-substituted and/or more complex substituents on aromatics.

Koch et al. (2007) also pointed out several liquid-phase studies that find evidence of ether bond cleavage. Again, *ipso* addition of hydroxyl radical is cited as being responsible for ether bond cleavage (removal of an oxy substituent) and C-C cleavage (removal of a methyl substituent) in some cases. Aihara et al. (1993) studied two pathways in which 1,2-dihydroxybenzene (catechol) forms from *o*-methoxyphenols using a copper (II) – ascorbic acid-dioxygen system: hydroxyl radical attack *ipso* to the methoxy containing carbon, and H-abstraction from the methoxy group. Hydroxyl radical attack *ipso* to the methoxy containing carbon allows for loss of the methoxy (–OCH<sub>3</sub>) group in the form of methanol (CH<sub>3</sub>OH), removing a C atom from the parent molecule. During H-abstraction from the methoxy group, Aihara et al. (1993) propose that the OCH<sub>2</sub> radical reacts with the OH radical and then elimination of formaldehyde (HCHO) ensues. In isotopically labeled studies, Aihara et al. (1993) observed transformation of guaiacol to catechol and syringol to hydroxy guaiacol. They also measure yields of these products that might imply *ipso* addition of OH being a more important route than previously considered compared to *ortho*- and *para*-OH addition to the ring. This route could serve as explanation of the observed hydroxy substitutions of methoxy groups in the current systems, but they cannot be verified without measurements of methanol and formaldehyde.

While we see ions indicating that methoxy groups may convert to hydroxyl groups, we do not see evidence supporting the addition of or reversible exchange of methoxy groups to the ring. Methoxy radicals may be formed from photolysis of the C–OCH<sub>3</sub> bond, if we assume that this bond energy can be proxied by the bond in dimethylether (CH<sub>3</sub>–OCH<sub>3</sub>). This bond energy, 351.9 kJ mole<sup>-1</sup> (Luo, 2007), translates to an approximate wavelength of 340 nm, exactly where the irradiance spectrum of the light source in our atmospheric chambers peaks. In the event that methoxy radicals are formed, though, it seems unlikely that methoxy radical reaction with the methoxyphenols results in methoxy addition to the ring or replacement of a hydroxy group, especially compared to rapid reaction with O<sub>2</sub>. Gomez et al. (2001) studied the reaction rate constants for methoxy radical with cyclohexane, cyclohexene, and 1,4-cyclohexadiene. It was found that methoxy radical reaction with these compounds is likely to result in preferential H abstraction with minor routes of addition to cyclohexene and 1,4-cyclohexadiene. This corroborates the absence of >C<sub>7</sub> species in the guaiacol case and >C<sub>8</sub> species in the syringol case.



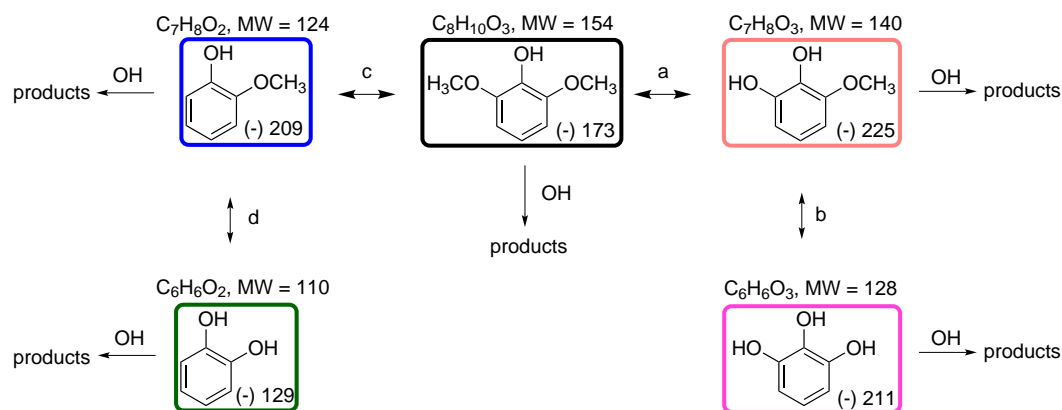


Fig. 8. Possible routes of methoxy loss in syringol gas-phase chemistry.

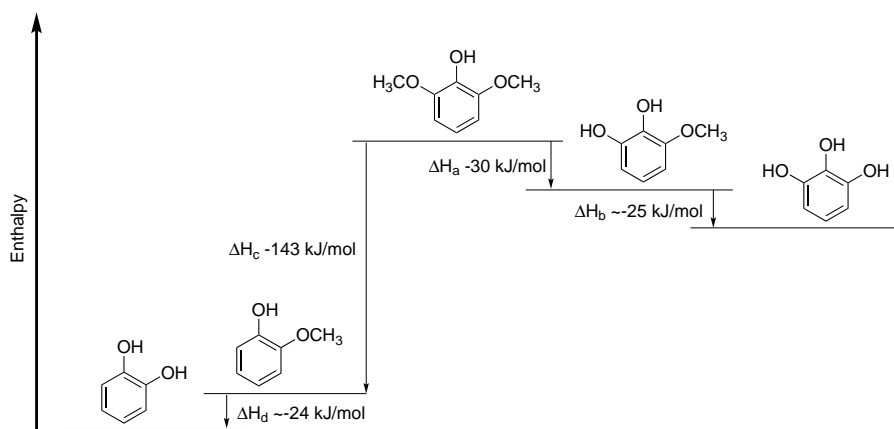


Fig. 9. Estimates of enthalpy differences between several products indicating methoxy loss.

Thermodynamic evaluation of the substituent group effects in methoxyphenols can also provide some insight. The following discussion is based on the work presented in Varfolomeev et al. (2010) that examines pairwise substitution effects, and inter- and intramolecular hydrogen bonds in methoxyphenols and dimethoxybenzenes. Varfolomeev et al. (2010) show that intramolecular hydrogen bonding plays a significant role on *ortho*-methoxy substituted phenols. This leads to *o*-methoxyphenols having a lower phenolic O–H bond dissociation energy (BDE) compared to *meta* isomers and higher than *para* isomers. This means that the phenolic O–H BDE increases in order the order 4-methoxyphenol < 2-methoxyphenol < 3-methoxyphenol. From the ab initio calculations of Varfolomeev et al. (2010), the enthalpy of formation for 2-methoxyphenol (guaiacol) is expressed below:

$$\Delta_f H_m^\circ(2\text{-methoxyphenol}) = \Delta_f H_m^\circ(\text{B}) + \Delta H(\text{H} \rightarrow \text{OH}) \\ + \Delta H(\text{H} \rightarrow \text{OCH}_3) + (\textit{ortho}\text{-OH-OCH}_3),$$

where  $\Delta_f H_m^\circ(\text{B})$  = enthalpy of formation of benzene,  $\Delta H(\text{H} \rightarrow \text{OH}) = -179.0 \text{ kJ mol}^{-1}$ , the increment in enthalpy for substitution of an H atom by  $-\text{OH}$ ,  $\Delta H(\text{H} \rightarrow$

$\text{OCH}_3) = -150.5 \text{ kJ mol}^{-1}$ , the increment in enthalpy for substitution of an H atom by  $-\text{OCH}_3$ , and (*ortho*-OH–OCH<sub>3</sub>) =  $4.5 \text{ kJ mol}^{-1}$ , the increment in enthalpy for the pairwise interactions of  $-\text{OH}$  and  $-\text{OCH}_3$  *ortho* to one another.

The enthalpy of formation expression begins with the base enthalpy of formation from benzene, then adds the effects of hydroxy or methoxy substitution to the ring. Finally, the pairwise interactions are added, reflecting the stabilizing intramolecular H bonding and destabilizing effect from sterics in the *ortho*-OH–OCH<sub>3</sub> configuration. From the comparison of the substitution effects alone, this suggests that H→OH is slightly more favorable than H→OCH<sub>3</sub>. For example, this might support substitution of a methoxy group by a hydroxy group in the guaiacol system, as the pairwise *ortho* effect is minimal.

Assuming that the Varfolomeev et al. (2010) formulation can be extended to the case of a trisubstituted benzene, we can write an analogous expression for the case of syringol (2,6-dimethoxyphenol):

$$\begin{aligned} \Delta_f H_m^\circ(2,6\text{-dimethoxyphenol}) &= \Delta_f H_m^\circ(\text{B}) + \Delta H(\text{H} \rightarrow \text{OH}) \\ &+ 2 \times \Delta H(\text{H} \rightarrow \text{OCH}_3) + 2 \times (\textit{ortho}\text{-OH-OCH}_3) \\ &+ (\textit{meta}\text{-OCH}_3\text{-OCH}_3), \end{aligned}$$

where (*meta*-OCH<sub>3</sub>-OCH<sub>3</sub>) = 0.1 kJ mol<sup>-1</sup> from Varfolomeev et al. (2010).

After writing a similar expression for 3-methoxybenzene-1,2-diol (hydroxyguaiaicol), the difference between the substituent effects for hydroxyguaiaicol and syringol is -30 kJ mol<sup>-1</sup>, in slight favor of hydroxyguaiaicol. The pairwise effect of (*ortho*-OH-OH) was not calculated by Varfolomeev et al. (2010), but assuming that it is less sterically hindered compared to (*ortho*-OH-OCH<sub>3</sub>) and is stabilized by intramolecular hydrogen bonding similar to 1,2-dimethoxybenzene, we assume it to be < 4.5 kJ mol<sup>-1</sup>. Due to the combined favorability of energetics for the substituent effects and the pairwise effects, this can help explain why some of the syringol rapidly converts to hydroxyguaiaicol and even trihydroxybenzene. A simplified schematic with estimated enthalpy differences using this analysis is presented in Fig. 9.

The “phenolic route”, involving OH addition at the *ortho* position to the main electron-donating substituent group has been cited as being the main channel responsible for SOA formation in many aromatic systems (Calvert et al., 2002; Johnson et al., 2005; Nakao et al., 2011). However, with the additional methoxy groups, OH attack at the alpha-unsubstituted carbons to the methoxy groups becomes competitive and changes the potential for maintaining aromaticity in the methoxyphenol systems. It is possible that the extra methoxy group in syringol inhibits OH attack that is typically favored *ortho* to the hydroxyl substituted carbon in other systems that generate SOA via this route. For the other compounds, in order to achieve the degree of oxidation of the aerosol, it appears that at least two steps of reaction are needed, but that these steps occur fairly rapidly. If subsequent OH reactions are even slower due to the extra methoxy group in the case of syringol, then conversion to SOA may not be as complete as with the other two compounds over the duration of the experiments reported here.

#### 4.6 Chemical basis for observed yields

The AMS high-resolution spectra for the methoxyphenol systems are distinct. Initial aerosol growth in the guaiacol system is marked by the exact mass ion C<sub>4</sub>H<sub>3</sub>O<sub>2</sub><sup>+</sup>. Later the aerosol growth is characterized by the ions C<sub>2</sub>H<sub>3</sub>O<sup>+</sup>, and C<sub>2</sub>H<sub>5</sub>O<sub>2</sub><sup>+</sup>. In the case of syringol, initial growth is characterized by C<sub>5</sub>H<sub>2</sub>O<sub>2</sub><sup>+</sup> and C<sub>7</sub>H<sub>9</sub>O<sub>3</sub><sup>+</sup> ions, followed by C<sub>5</sub>H<sub>5</sub>O<sub>4</sub><sup>+</sup>, and then by C<sub>5</sub>H<sub>2</sub>O<sub>4</sub><sup>+</sup> and C<sub>2</sub>H<sub>4</sub>O<sub>3</sub><sup>+</sup>. This suggests that guaiacol SOA is characterized by smaller carbon number and less oxygenated fragments than the case of syringol.

The characteristic ions from the AMS spectra for syringol SOA seem counter-intuitive for understanding why the sy-

ringol SOA yield is lower than that of guaiacol. Syringol has more carbon and oxygen to begin with and these ions are larger, more oxygenated fragments than in guaiacol. However, if we return to the gas-phase comparison of the characteristic trends in the guaiacol (Fig. 5) and syringol (Fig. 7) systems, one notes that many of the major products from syringol photooxidation peak later than their analogs in the guaiacol system. Further, there is no analogous peak in the guaiacol system to the *m/z* 275 unique to the syringol system. Since this is a major product of gas-phase syringol photooxidation and it peaks > 4 × 10<sup>6</sup> molec cm<sup>-3</sup> h of OH exposure compared to guaiacol ions that peak generally within 3 × 10<sup>6</sup> molec cm<sup>-3</sup> h, it acts like a gas-phase carbon reservoir. That is, the signal at *m/z* 275 is much more significant than the more efficient SOA formers at *m/z* 129 (dihydroxybenzene, not shown) and *m/z* 209 (guaiaicol) in Fig. 7b.

AMS elemental analyses of SOA from these systems as reported in Chhabra et al. (2011) suggest that the majority of oxygenation is derived from organic acid functionalities: 77 % for phenol, 61 % for guaiacol, and 59 % for syringol. This is consistent with the necessity of further oxidation of ring-opened dicarbonyl fragments and ring-opened acids to explain the presence of highly oxidized species observed in the guaiacol filter analyses. Ofner et al. (2010) found that the SOA-formation process from catechol involves ring-opened carboxylic acid derivatives that are part of highly hydroxylated olefin structures, which supports the organic acid functionality observed in the current related systems. Previous study on the ozonolysis of 1,2-dihydroxybenzene showed that several carbonyls existed in both the gas and particle phases (Olariu, 2001). Carboxylic acids believed to be ring fragments, as well as muconic acid, a ring-opened C<sub>6</sub> dicarboxylic acid, were also identified. Muconic acid was also identified in the particle phase from benzenediol ozonolysis (Tomas et al., 2003). These results support the current observations that many dicarbonyl species in general trend with the SOA mass, suggesting that they may be semi-volatile in nature.

The decrease in oxygenation derived from organic acid functionalities from phenol to the methoxyphenol systems may reflect the oxygen reserved in OH adduct and aromatic retaining pathways, especially in the case of syringol (Fig. 8). This may also partially explain why the syringol SOA yield is lower if particle-phase acid formation from carbonyl ring fragments drives the SOA growth. At similar OH exposures, the syringol system still has not developed a carbonyl ring fragment signal that tracks well with the SOA growth, as *m/z* 185 does in the phenol system (Fig. 2) and *m/z* 149 does in the guaiacol system (Fig. 5). The *m/z* 245 fragment lags the SOA growth in the syringol system (Fig. 7a). This could be due to the lack of OH addition to the ring on syringol in critical positions that lead to SOA formation via ring-fragmentation pathways and acid formation.

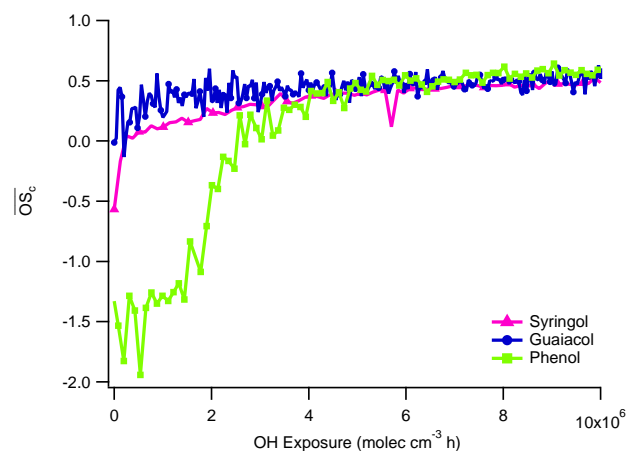
Estimation of carbon oxidation state also provides some insight into the chemical evolution of the aerosol towards

incorporation of highly oxidized species. Representative time series of the carbon oxidation state,  $\overline{\text{OS}}_{\text{C}}$ , for the studied systems are shown in Fig. 10. Carbon oxidation state was estimated with the AMS O:C and H:C elemental analyses using the formulation presented in Kroll et al. (2011). All systems approach an  $\overline{\text{OS}}_{\text{C}}$  of +0.5. This approaches the range of low-volatility oxygenated organic aerosol (Kroll et al., 2011). A delay in the rise of  $\overline{\text{OS}}_{\text{C}}$  from phenol aerosol is observed, consistent with the rise in organic growth at the point of formation of trihydroxybenzene (Fig. 2). Syringol aerosol lags in oxidation in the beginning compared to guaiacol until approximately  $4 \times 10^6 \text{ molec cm}^{-3} \text{ h}$  of OH exposure. This times with the peak of many of the methoxy loss products in Fig. 7b, supporting that additional oxidation is achieved once methoxy loss occurs and the syringol product distribution overlaps that of guaiacol (i.e. G + OH and THB) and the “phenolic route”. Primary organic aerosol from biomass burning was found to have an  $\overline{\text{OS}}_{\text{C}}$  of  $-1.0$  to  $-0.7$  (Aiken et al., 2008; Kroll et al., 2011), which falls inbetween the range of  $\overline{\text{OS}}_{\text{C}}$  from the three systems here. However, aged (Whistler Mountain) aerosols ranged  $-0.6$  to  $+0.6$  (Sun et al., 2009; Kroll et al., 2011), enveloping the methoxyphenols and a portion of the phenol  $\overline{\text{OS}}_{\text{C}}$  evolution. In terms of average carbon oxidation state, these experiments pick up from primary organic aerosol in biomass burning and move towards the character of aged organic aerosol.

## 5 Atmospheric relevance

While the current experiments were conducted without the addition of  $\text{NO}_x$ , the chemistry elucidated under these low- $\text{NO}_x$  conditions is expected to be relevant for typical atmospheric conditions. For phenol photooxidation in the presence of  $\text{NO}_x$ , Olariu et al. (2002) measured yields of 0.8 catechol (1,2-dihydroxybenzene), 0.037 1,4-benzoquinone, and 0.058 2-nitrophenol. However, according to kinetic data and yield comparisons for nitrophenol presented in Berndt and Boge (2003), it is likely that this nitro-product formation can be biased due to sufficiently high  $\text{NO}_2$  ( $\lesssim 800$  ppb) concentrations. Berndt and Boge (2003) point out that for atmospheric conditions with  $\text{NO}_2$  of approximately 20 ppb, it is probable that phenoxy radicals also react with  $\text{O}_3$  as well as  $\text{NO}$  and  $\text{NO}_2$ . This would result in more minor nitrophenol yields in the atmosphere for daytime chemistry. Reaction with  $\text{NO}_3$  should also be considered as phenolic compounds can react rapidly with trace levels of  $\text{NO}_3$  during the day and especially at night.

Thus, because OH addition dominates over H abstraction, and subsequent  $\text{O}_2$  addition to the OH radical adduct is rapid, it is more likely that in the atmosphere, SOA primarily forms via pathways involving higher-generation OH adducts and their respective routes to oxygenated fragments. This is consistent with the chemistry presented in this work. Though we did not observe appreciable evidence of gas-phase hydroper-



**Fig. 10.** Evolution of the average carbon oxidation state ( $\overline{\text{OS}}_{\text{C}}$ ) for studied systems.

oxides, they may be rapidly incorporated into the particle phase and contribute to peroxide formation, as postulated by Johnson et al. (2005) and observed by Sato et al. (2007) for the case of toluene photooxidation. Sato et al. (2007) also find similar species in the particle-phase under  $[\text{NO}]_0 = 0.2$  and 1 ppm, still citing  $\text{O}_2$  addition to the OH-aromatic adduct as the major channel contributing to SOA formation.

In the presence of moderate  $\text{NO}_x$ , we would expect the major gas-phase species to be the same as those presented here, with only a minor contribution from nitrogen-containing products. Still, for the studied compounds, novel routes of reaction of alkylperoxy radicals under these low- $\text{NO}_x$  conditions may still result in product distributions expected under higher  $\text{NO}_x$  conditions. First, the aromatic peroxy radicals in the studied systems may preferentially decompose to regenerate OH via novel  $\text{RO}_2 + \text{HO}_2$  chemistry (Birdsall et al., 2010). This is a possible explanation for the observation of ring-opened acids in phenol and guaiacol. Second, generation of carbonyls via novel rearrangements of ether peroxy radicals (Orlando and Tyndall, 2012), relevant for guaiacol and syringol, would further make the product distributions less distinct from alkoxy radical pathways expected under high- $\text{NO}_x$  chemistry.

These observations warrant caution when conducting experiments in the presence of  $\text{NO}_x$  for aromatic compounds, as derived aerosol yields may be based on chemistry non-representative of the atmosphere. Hence, we left discussion of the experiments conducted in the presence of  $\text{NO}_x$  for the Appendix. Klotz et al. (2002) found that at high- $\text{NO}_x$  conditions (hundreds of ppb to ppm) one must also start to consider reactions of  $\text{NO}_x$  with the hydroxy radical adduct. At sufficiently high  $\text{NO}_x$  concentrations, the gas-phase chemistry can generate higher yields of nitroaromatics as compared to the OH aromatic adduct (Koch et al., 2007; Nishino et al., 2008). Nitroaromatics can serve to enhance aerosol yields if sufficient nitro-containing functionalities are achieved or

to act as gas-phase carbon reservoirs and depress aerosol yields. Hydroxyl groups will lead to a vapor pressure lowering more than that associated with nitro groups, but sufficient nitro groups could lead to incorporation to the particle phase. Owing to the lower OH reactivity of the molecule containing a nitro group as compared to the comparable molecule containing a hydroxyl group (Kwok and Atkinson, 1995), one might also find that more OH is necessary in the presence of  $\text{NO}_x$  to attain further gas-phase development to more highly oxidized species that contribute to SOA formation.

While the OH-initiated photooxidation of aromatics in general is complex (Calvert et al., 2002), methoxyphenols seem to follow some steps of the general mechanism developed for simpler aromatics. We have qualitatively presented gas-phase trends that give some insight into the chemistry of SOA formation. Though, more complete mechanism development would be enhanced by the availability of authentic standards to solidify mass assignments and more detailed structural information. For example, the structural identity of the ion at  $m/z$  275 in the syringol system remains elusive. Calculation of quantitative gas-phase yields would benefit from additional kinetic data pertaining to some of the complex intermediates.

Carbon loss via the methoxy groups in guaiacol and syringol appears to be a feasible route to describe the gas-phase product distributions. While the chemistry of syringol is complex, this compound may be a better model system to represent the mixture of methoxy and phenolic compounds since it displays chemistry from both classes. Syringol, however, does not generate as much SOA as these other systems. This may be a result of the unique  $\text{OCH}_3\text{--OH--OCH}_3$  pairwise effects that hinder the typical phenolic pathway responsible for SOA formation.

Future studies to investigate this effect might start with methoxyphenols of different substituted positions relative to the hydroxy group to see the effect on SOA yields. For example, using a methoxyphenol where at least one or two methoxy groups are not *ortho* to the hydroxy group would serve to further investigate the importance of OH addition *ortho* to the hydroxy group in these compounds, as 1,2-dihydroxybenzene is the major gas-phase product and SOA former from phenol. Also, 2,5-dimethoxyphenol might also indicate if OH attack is preferred at the 6-position, again *ortho* to the hydroxy group. Such studies would require more detailed analyses that provide greater structural analysis for the gas-phase products as well as unique tracers for fragmentation patterns that would clarify the chemistry.

Detailed study of the gas-phase products from the OH-initiated photooxidation of methoxybenzene might also be informative for understanding the proposed carbon loss associated with methoxy groups in the guaiacol and syringol systems. If phenol is a major product, this has implications on understanding the chemical fate of methoxyphenol-like compounds in the atmosphere. For example, Lauraguais et al. (2012) suggest that syringol is too reactive with OH (1.8 h) to

be a relevant tracer in the atmosphere for woodsmoke emissions and that it results in a very minor SOA yield. However, the chemical analyses presented in this study suggest that syringol can efficiently convert to guaiacol and hydroxylated benzenes that are longer lived in the atmosphere and potentially have larger SOA yields. This can be analogous to the isoprene photochemical cascade, where it is the products of isoprene photooxidation that have greater SOA forming potential (Kroll et al., 2006). The syringol conversion to guaiacol also has implications for using these compounds as specific tracers for fuel type (hardwood vs. softwood). Thus, care must be taken in selecting one or more surrogate species for use in modeling aerosol yields from biomass burning emissions in the atmosphere.

## Appendix A

### Methoxyphenol oxidation in the presence of $\text{NO}_x$

We present here results from the methoxyphenol experiments performed in the presence of  $\text{NO}_x$  including SOA yields and qualitative analyses of the chemical development in these systems. As mentioned in the Introduction section, the main text of this paper focuses on methoxyphenol and phenol photooxidation under idealized chemical conditions ( $\text{NO}_x < 10$  ppb), such that the predominant fate of the  $\text{RO}_2$  radical here is reaction with  $\text{HO}_2$ . Biomass burning events certainly generate  $\text{NO}_x$ , and the authors recognize that the low- $\text{NO}_x$  chemistry presented in the main text by no means provides the complete chemical picture of aged biomass burning plumes. A complementary set of experiments in the presence of  $\text{NO}_x$  was conducted along with the low- $\text{NO}_x$  experiments presented earlier using the same protocols, except that HONO was used as the OH precursor instead of  $\text{H}_2\text{O}_2$  and NO was also injected.

However, we do not focus on the results of these experiments in the main text for both technical and scientific reasons including: (1) The actual experimental conditions resulted in hundreds of ppb levels of  $\text{NO}_x$ , which are much higher than those reported for biomass burning events. For example, Iinuma et al. (2010) measured biomass burning co-located  $\text{NO}_x$  on the order of 20 ppb. Thus, the product distributions in these experiments result in favored nitro-products, but do not necessarily represent the more relevant chemical development in the atmosphere.  $\text{NO}_x$  levels can also become depleted several hours later (via  $\text{NO}_2$  incorporation into PAN species) or with dilution via transport of a plume as observed in Akagi et al. (2012). Thus, experiments in the absence of or at levels of  $\text{NO}_x < 10$  ppb (such as those presented in the main text) can still be representative of product distributions at lower  $\text{NO}_x$  levels ( $< 30$  ppb) when methoxyphenol precursors are transported away and subsequently photooxidize. (2) Interference of HONO response on our  $\text{NO}_x$  analyzer made it more difficult to report

**Table A1.** Experimental conditions and aerosol yields in the presence of added NO<sub>x</sub>.

Date (2010)	Organic	[Org] <sub>0</sub> (ppb)	[Org] <sub>f</sub> (ppb)	RH <sub>0</sub> (%)	[NO <sub>2</sub> ] <sub>0</sub> <sup>a</sup> (ppb)	[NO] <sub>0</sub> <sup>a</sup> (ppb)	Vol <sub>0</sub> <sup>b</sup> μm <sup>3</sup> cm <sup>-3</sup>	Vol <sub>f</sub> <sup>b</sup> μm <sup>3</sup> cm <sup>-3</sup>	Yield
2/3	guaiacol	14.2	0.7	< LDL <sup>c</sup>	447	282	12.9	27.5	0.34 ± 0.02
2/5	guaiacol	51.9	3.1	5.6	536	461	13.5	92.5	0.53 ± 0.01
3/2	guaiacol	38.0	2.0	< LDL <sup>c</sup>	458	295	10.5	65.4	0.46 ± 0.01
3/28	guaiacol	13.6	0.8	< LDL <sup>c</sup>	429	222	11.4	26.5	0.35 ± 0.02
3/31	guaiacol	55.3	4.0	< LDL <sup>c</sup>	518	330	13.8	86.5	0.46 ± 0.02
4/6	guaiacol	16.4	1.2	< LDL <sup>c</sup>	493	314	9.0	29.4	0.4 ± 0.02
4/9 <sup>d</sup>	guaiacol	14.2	0.5	< LDL <sup>c</sup>	509	315	18.7	33.5	0.35 ± 0.02
4/13 <sup>d</sup>	guaiacol	15.4	0.8	< LDL <sup>c</sup>	464	296	11.3	27.1	0.4 ± 0.02
4/15	guaiacol	35.6	3.7	< LDL <sup>c</sup>	526	321	10.8	55.7	0.49 ± 0.02
4/22	guaiacol	50.4	4.4	< LDL <sup>c</sup>	433	329	14.4	74.6	0.47 ± 0.02
2/19	phenol	14.8	1.9	< LDL <sup>c</sup>	443	295	14.0	32.0	0.54 ± 0.06
3/23	phenol	23.5	3.0	< LDL <sup>c</sup>	488	301	14.2	37.9	0.54 ± 0.06
3/6	syringol	53.1	2.4	< LDL <sup>c</sup>	448	286	11.9	68.0	0.3 ± 0.01
3/9	syringol	124.2	4.2	< LDL <sup>c</sup>	481	313	14.6	110.0	0.2 ± 0.01
3/11	syringol	92.3	2.9	< LDL <sup>c</sup>	469	314	12.6	73.0	0.17 ± 0.01
3/13	syringol	31.9	1.0	< LDL <sup>c</sup>	418	241	12.8	29.0	0.11 ± 0.01

<sup>a</sup> Corrections for HONO interference not made; values for NO<sub>x</sub> listed are upper limits

<sup>b</sup> Vol<sub>0</sub> is initial seed volume; Vol<sub>f</sub> is final volume (seed + organic aerosol), as measured by DMA

<sup>c</sup> RH probe LDL = 5 % RH

<sup>d</sup> Additional HONO injected mid-experiment

precise NO<sub>x</sub> levels. (3) Before photooxidation commenced, we observed a likely nitro-product that formed once HONO was introduced to the reactors, presumably a reaction of methoxyphenol with HONO and/or NO<sub>x</sub> (by-product of the HONO synthesis) in the dark to form a nitro-product. This was not observed in Lauraguais et al. (2012) for the case of syringol and CH<sub>3</sub>ONO in the dark. While direct HONO/NO<sub>x</sub> reaction with the methoxyphenol compounds may be of atmospheric relevance, this complicated the chemical system that we started with and made it difficult to constrain the chemistry for the observed product distributions.

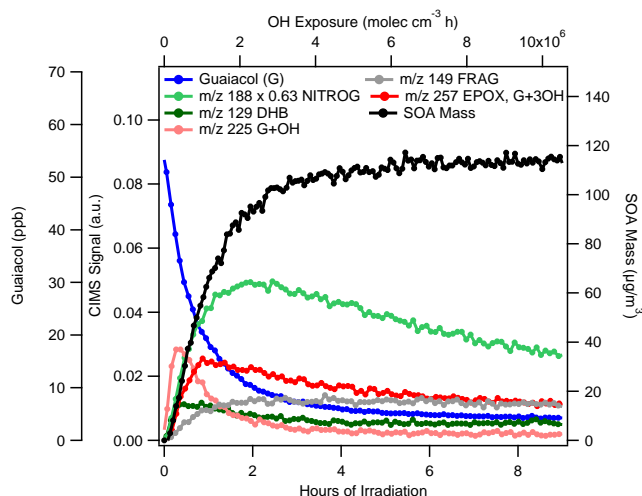
#### A1 Methoxyphenol SOA yields in the presence of NO<sub>x</sub>

The yield data presented here are not recommended for use in atmospheric models, as the conditions of the experiments are not representative of typical or biomass burning NO<sub>x</sub> levels in the atmosphere. The excess of NO<sub>x</sub> employed in these experiments likely biases the product distributions to more nitrogen-containing species, thereby affecting associated volatilities and the SOA yields. All experiment conditions and SOA yields are presented in Table A1.

For high-NO<sub>x</sub> experiments (NO > 200 ppb initially), nitrous acid (HONO) was used as the OH precursor. HONO was prepared by adding 15 mL of 1 wt % aqueous NaNO<sub>2</sub> dropwise into 30 mL of 10 wt % sulfuric acid in a glass bulb. A stream of dry air was then passed through the bulb, sending HONO into the chamber. During this process NO and NO<sub>2</sub> formed as side products and were also introduced into the

chamber. NO/NO<sub>x</sub> was measured with a commercial chemiluminescence NO<sub>x</sub> monitor (Horiba, APNA-360) with lower detection limit (LDL) of 5 ppb of NO. Total starting NO<sub>x</sub> was generally just under 800 ppb as indicated by the NO<sub>x</sub> analyzer, and no corrections have been made for HONO interference on the analyzer.

Additional HONO was injected in two ~ 14 ppb guaiacol experiments (4/9 and 4/13) to explore the extent to which additional OH and NO<sub>x</sub> would affect the gas-phase chemistry leading to SOA formation. Additional HONO was injected both at the mid-point of aerosol growth (4/13) and at the point where virtually all of the initial organic had been consumed (4/9). In both cases, further oxidation by OH did not lead to further aerosol formation in comparison to the guaiacol high-NO<sub>x</sub> experiment with similar organic loading (2/3). While an enhancement of some nitrogen-containing products occurred as observed in the gas phase for the cases with additional HONO (4/9 and 4/13), this suggests that the major components of SOA formation are not likely from these particular products. In addition, there were no observations of new species in the CIMS spectra to suggest that additional OH exposure unlocks new generations of chemistry that contribute to SOA formation. These observations are consistent with the relatively rapid formation of highly oxidized, low-volatility SOA during the course of the experiment.



**Fig. A1.** Guaiacol high- $\text{NO}_x$  gas-phase and particle-phase development.

## A2 Phenol high- $\text{NO}_x$ chemistry

Olariu et al. (2002) measured yields of 0.8 catechol (1,2-dihydroxybenzene), 0.037 1,4-benzoquinone, and 0.058 2-nitrophenol using Fourier transform infrared (FT-IR) absorption spectroscopy for phenol photooxidation in the presence of  $\text{NO}_x$ . The CIMS traces in this study are consistent with these products. We find the major signals in the CIMS at  $m/z$  129 (dihydroxybenzene) and  $m/z$  158 (nitrophenol). There is some signal at  $m/z$  211 and  $m/z$  174 indicating that trihydroxybenzene and its nitro-product are present in small amounts. However, according to the kinetic data and yield comparisons for nitrophenol presented in Berndt and Boge (2003), it is likely that nitro-product formation is favored from reaction of phenoxy radicals with  $\text{NO}_2$  due to sufficiently high  $\text{NO}_2$  concentrations in these experiments (max.  $\text{NO}_2 \sim 500$  ppb). The relative signals in the CIMS indicated such a bias existed. Berndt and Boge (2003) point out that for atmospheric conditions with  $\text{NO}_2$  of approximately 20 ppb, it is probable that phenoxy radicals also react with  $\text{O}_3$  and other species. In negative mode operation, the CIMS is not sensitive to quinone products, however, there are masses in positive mode operation at  $m/z$  109 and 127 that may be the  $[\text{M}+1]^+$  and  $[\text{M}+19]^+$  ions of benzoquinone. The set of  $\text{C}_6$  ring-opened acids as proposed in Fig. 3 for low- $\text{NO}_x$  conditions is not prevalent under high- $\text{NO}_x$  conditions, except for the acid at  $m/z$  145. An alkoxy radical generated from  $\text{RO}_2 + \text{NO}$  results in the same acid ( $\text{C}_6\text{H}_6\text{O}_3$ ). Signals for proposed acidic fragments at  $m/z$  185 and  $m/z$  201 in the low- $\text{NO}_x$  case are also present under high- $\text{NO}_x$  conditions.

## A3 Guaiacol high- $\text{NO}_x$ chemistry

During experiments with HONO injection, the majority of guaiacol is rapidly converted to nitroguaiacol ( $\text{C}_7\text{H}_7\text{NO}_4$ )

**Table A2.** SOA products observed in UPLC/(-)ESI-ToFMS offline filter analysis for the guaiacol high- $\text{NO}_x$  system.

$[\text{M}-\text{H}]^-$	Meas. Mass	Suggested Chemical Formula (M)	Error* (mDa)
154	154.0121	$\text{C}_6\text{H}_5\text{NO}_4$	-1.9
170	170.0069	$\text{C}_6\text{H}_5\text{NO}_5$	-2
186	186.0023	$\text{C}_6\text{H}_5\text{NO}_6$	-1.6
199	198.9965	$\text{C}_6\text{H}_4\text{N}_2\text{O}_6$	-2.6
200	200.0182	$\text{C}_7\text{H}_7\text{NO}_6$	-1.3
184	184.0229	$\text{C}_7\text{H}_7\text{NO}_5$	-1.7

\* The error in the accurate mass measurements is the difference between the theoretical mass of the suggested molecular formulae and the measured mass of the ion.

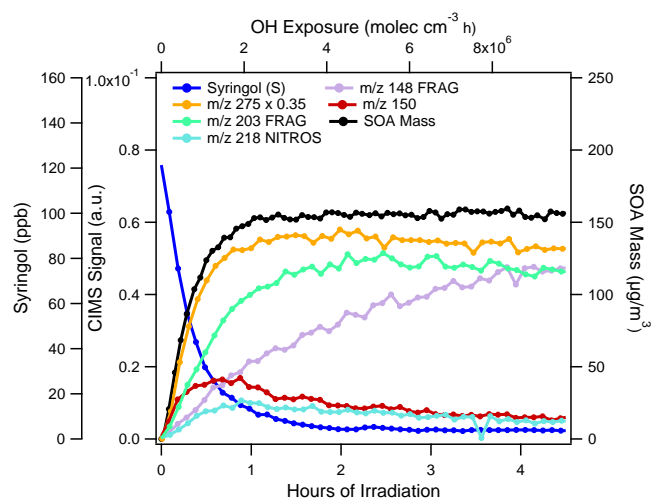
monitored at  $m/z$  188, NITROG, in Fig. A1. Production of dihydroxybenzene and the OH adducts still occurs as seen in the low- $\text{NO}_x$  case, as well as their nitro analogs in the high- $\text{NO}_x$  case. However, the nitro OH adducts are formed in minor amounts. The hydroperoxide channel is negligible as expected. However, the presence of  $\text{NO}_x$  seems to favor the production of many nitrogen-containing fragments since there is a prevalence of even-numbered  $m/z$ 's  $< 200$  in the CIMS spectra. Some of the identifications of these ions remains elusive. Additional even-numbered fragments are present, but of low signal to noise.

Table A2 lists a few of the most prevalent particle-phase products (from highest to lowest peak area). The filter data show that dihydroxybenzene with a nitro group ( $\text{C}_6\text{H}_5\text{NO}_4$ ) is incorporated into the particle phase as well as up to two more oxygens ( $\text{C}_6\text{H}_5\text{NO}_5$ ,  $\text{C}_6\text{H}_5\text{NO}_6$ ). The formation of these essentially non-volatile products and rapid incorporation into the particle phase is consistent with the lack of CIMS observations for such oxygenated compounds in the guaiacol high- $\text{NO}_x$  system. The presence of particle-phase  $\text{C}_6\text{H}_4\text{N}_2\text{O}_6$  suggests that this product may be a  $\text{C}_6$  ring fragment that undergoes additional oxidation. The  $\text{C}_7$  products are consistent with the minimally detected gas-phase formation of nitroguaiacol OH adducts, also supporting rapid incorporation into the particle phase. One thing to note is that while the filter data suggest that the majority of particle-phase products are nitrogen containing, this may be partially biased by sensitivity of the technique for measuring such compounds. The AMS bulk analyses show that the guaiacol system N : C ratio is only 0.04 which does not match the average N : C represented by the products in Table A2. However, the AMS N : C ratio may also be biased low by uncertainties introduced if many of these species are organonitrates (Farmer et al., 2010).

## A4 Syringol high- $\text{NO}_x$ chemistry

High- $\text{NO}_x$  syringol photooxidation is dominated primarily by  $m/z$  275 again, and fragments (Fig. A2). The ion at  $m/z$  203 is likely a volatile fragment and was also prominent





**Fig. A2.** Syringol high- $\text{NO}_x$  gas-phase and particle-phase development.

in the guaiacol high- $\text{NO}_x$  system. Nitrosyringol is a small contribution at  $m/z$  218 (NITROS) and unidentified  $m/z$  150 grows in and reacts quickly in a similar manner. An ion at  $m/z$  148 growing in over time indicates the conversion of some of these initial nitro-containing products to nitrogen-containing fragments. This may also explain why the aerosol yields in the high- $\text{NO}_x$  case appear to be slightly lower than those in the low- $\text{NO}_x$  case. Particle-phase analyses and structure identification would be necessary to further understand if the particle phase consists mostly of  $< \text{C}_8$  species.

While in the low- $\text{NO}_x$  case one is able to distinguish some conversion of syringol to guaiacol and hydroxy guaiacol, the presence of  $\text{NO}_x$  seems to rapidly lead to incorporation of a nitro group. Nitroguaiacol species are minor species in this system. The SOA forms quickly in this case and could be due to rapid incorporation of the nitroguaiacol and nitrosyringol products into the particle phase.

#### A5 Chemical differences between high- $\text{NO}_x$ and low- $\text{NO}_x$ chemistry

In the case of guaiacol and syringol, product identifications are generally analogous for the low and high- $\text{NO}_x$  cases, except that more nitrogen-containing species are present in the high- $\text{NO}_x$  case. Both cases result in fragments that are often acidic and have the potential to become further oxidized. It is likely that these highly oxidized fragments are incorporated into the particle phase as observed in the filter analyses for the guaiacol low- $\text{NO}_x$  system. If alkylperoxy radicals preferentially decompose to regenerate OH via novel  $\text{RO}_2 + \text{HO}_2$  pathways as observed from aromatic peroxy radicals (Birdsall et al., 2010) or to generate carbonyls via novel rearrangements of ether peroxy radicals (Orlando and Tyndall, 2012), this further makes the product distributions less distinct from alkoxy radical pathways in high- $\text{NO}_x$  chemistry.

In the guaiacol system, the acidic fragment traces in the CIMS have the same time profile as the SOA growth in the low- $\text{NO}_x$  case. Under high- $\text{NO}_x$  conditions, the SOA growth does not follow the same time profile as any of the major gas-phase products. This suggests that under the experimental conditions of OH exposure, the SOA under low- $\text{NO}_x$  conditions may be dominated by production of more highly oxidized and lower volatility products than under high- $\text{NO}_x$  conditions. The major difference between SOA formation in the low- and high- $\text{NO}_x$  conditions may be attributed to the nitro-group incorporation into the ring. Comparison of the predominant first-generation product nitroguaiacol ( $\text{C}_7\text{H}_7\text{NO}_4$ ) under high- $\text{NO}_x$  conditions to any of the  $\text{C}_7$ -retaining first generation products (typically  $\text{O} > 5$ ) under low- $\text{NO}_x$  conditions demonstrates that the low- $\text{NO}_x$  case begins with relatively lower vapor pressure products. Though we did not observe appreciable evidence of gas-phase hydroperoxides in the low- $\text{NO}_x$  case, they may be rapidly incorporated into the particle phase. With continued OH addition in the low- $\text{NO}_x$  case and competing nitro-group addition in the high- $\text{NO}_x$  case, the volatilities resulting from addition of the hydroxyl and nitro groups can rival one another. Though, hydroxyl groups will lead to a vapor pressure lowering more than that associated with nitro groups. In addition, the nitroguaiacol fate is predominantly gas-phase reaction with OH, as it does not have sufficient oxygens as represented by the majority of products observed in the filter data ( $> \text{O}_4$ ). Owing to the lower OH reactivity of the molecule containing a nitro group as compared to the comparable molecule containing a hydroxyl group (Kwok and Atkinson, 1995), this also suggests that more OH is necessary in the high- $\text{NO}_x$  case to attain further gas-phase development to more highly oxidized species. For example, while  $\text{RO}_2 + \text{NO}$  chemistry generally leads to fragmentation, there is a tendency for at least  $\text{C}_6$  highly oxygenated nitrogen containing species to be retained. This could be due to lack of continued OH exposure. It is also possible that the  $\text{C}_6$  species are derived from the dihydroxybenzene route in guaiacol photooxidation, rather than  $\text{RO}_2 + \text{NO}$  induced carbon-carbon cleavage.

In general, the CIMS spectra indicate there is greater signal at the fragmented acidic species compared to the ring retaining species, and these fragment traces tend to have a time profile that is in equilibrium with SOA growth. Though there are distinct nitrogen-containing products in the high- $\text{NO}_x$  case, it is again the fragments (many at the same  $m/z$  as the low- $\text{NO}_x$  case) and some nitrogen-containing fragment traces that seem to be in equilibrium with the particle phase. Thus, while nitrogen incorporation may occur early on in the high- $\text{NO}_x$  cases, there is still a degree of oxygenation required for these species to be incorporated in the particle phase. This is consistent with the observation of several aerosol species in the guaiacol high- $\text{NO}_x$  system that have chemical formulae with one nitrogen, but  $> 3$  oxygens (more than guaiacol itself). AMS elemental analyses of SOA from these systems as reported in Chhabra et al. (2011) suggest



that the majority of oxygenation is derived from organic acid functionalities under both NO<sub>x</sub> conditions, ranging 61–77 % for phenol and guaiacol and 54–59 % for syringol. There is a < 7 % contribution from organic nitrate groups under high-NO<sub>x</sub> conditions. The relatively lower contribution of oxygenation from organic acid groups in the syringol system compared to phenol and guaiacol may partially explain why the syringol SOA yield is also lower. SOA formation from these compounds may be dominated by acidic species that result from ring-opening pathways, so prevention of this pathway (as evidenced by the presence of *m/z* 275 under both conditions of NO<sub>x</sub>) in the case of syringol would hinder SOA formation.

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